

# The grand challenges in carbon capture, utilization, and storage

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## INTRODUCTION

If we were in a global war against climate change, we would carry out large-scale carbon capture, utilization, and storage (CCUS) (Smit et al., 2014). Some argue that if we fight the war against climate change via CCUS, this implies that we are promoting the continued use of fossil fuels instead of replacing fossil fuels by renewable energy such as solar and wind. At present, the contribution of fossil fuels in our energy supply is over 80%, while the renewable is only 10% (International Energy Agency (IEA), 2013). Most, if not all, of the energy scenarios predict an increase in the share of the renewables, but in absolute numbers the fossil fuels will continue to provide most of our energy needs in the foreseeable future. This is simply because the growth in renewable energy will not be able to keep up with our increasing energy demand associated with a growing world population. In such a scenario, a war against climate change without CCUS implies that we have to dramatically reduce our current energy consumption, and hence, accept a dramatic reduction in GDP. Thus, the adaptation of large-scale CCUS might be inevitable to mitigate ever-increasing CO<sub>2</sub> emissions with given population growth predictions.

At present, there are still very few signs of starting a war against climate change soon. The consequence is that we may significantly overshoot CO<sub>2</sub> levels in the atmosphere before any serious action is taken. In such a scenario, it is very likely that we also need to deploy technologies that can achieve negative emissions (i.e., direct CO<sub>2</sub> capture from air). In this scenario, the price of carbon will be so

high that any technology that uses CO<sub>2</sub> as a source of carbon will have such an economic advantage that CO<sub>2</sub> will replace fossil fuels for those applications (e.g., plastics and soaps). Thus, CO<sub>2</sub> utilization next to storage will be an integral component of carbon management.

## CARBON CAPTURE

For carbon capture (Wilcox, 2012), it is important to distinguish between stationary sources (power plants, factories, etc.) and mobile sources (cars, airplanes, etc.) of CO<sub>2</sub>. At present, there are no practical solutions for on-board capture CO<sub>2</sub> directly from mobile sources; therefore, we focus in the remainder on capturing CO<sub>2</sub> from stationary sources.

## COAL-FIRED POWER PLANTS

Removing CO<sub>2</sub> from gases emitted from stationary sources can be done, using quite old technology. Most natural-gas contains more CO<sub>2</sub> than is allowed to put in pipelines. Hence, gas companies use the amine scrubbing process developed by Bottoms (1930) to separate CO<sub>2</sub> from methane. A very similar process can be used to remove CO<sub>2</sub> from flue gas. There is very little doubt in the engineering community that this amine absorption technology can be scaled-up and implemented to capture flue gases. The problem, however, is that the regeneration of the amine solution and the subsequent compression of CO<sub>2</sub> for transport and geological storage is very energy intensive. As a consequence, a power plant with carbon capture will not only be more expensive to build, but also will have reduction in efficiency as high as 35% (Herzog et al., 1993). Research is therefore focused

on increasing the efficiency of the absorption process and on finding alternatives (e.g., solid adsorption or membranes).

Building a new power plant in which carbon capture would be added from the very beginning would give more possibilities to optimize the efficiency of combined power generation and carbon capture. One such example is oxy-combustion in which coal is burned with pure oxygen, and CO<sub>2</sub> is captured simply by condensing the water. In oxy-combustion, however, one needs to separate oxygen from the air, which is also an energy intensive separation. Other alternatives include, chemical looping, and integrated gasification combined cycle (IGCC), in which coal is converted into syngas, and the carbon capture process involves separation from H<sub>2</sub>.

## CARBON CAPTURE FROM DILUTE SOURCES

Thermodynamics tells us that the lower the concentration of CO<sub>2</sub>, the more expensive it is to capture a ton of CO<sub>2</sub> (House et al., 2011; Wilcox, 2012). Hence, most carbon capture effort has been initially focused on flue gas streams from coal-fired power plants, which contain about 15% CO<sub>2</sub>. As coal was the cheapest source of fossil fuel for many years, such an objective made perfect sense. However, the scenarios did not forecast the availability of large amounts of cheap natural gas in recent years. Replacing coal-fired power plants by gas-fired power plants will significantly reduce CO<sub>2</sub> emissions, but the scenario of burning all natural-gas reserves without CCS is only marginally less disastrous compared to burning all coal without CCS.

In the recent IPCC report (Stocker et al., 2013), the importance of negative

emissions has been established. According to the most recent IPCC report, however, we are currently emitting more CO<sub>2</sub> than predicted by the most pessimistic IPCC scenario. Therefore, it is more than likely that we will overshoot the target CO<sub>2</sub> levels. In such a scenario, one may need technologies to actively mitigate the effects of climate change (Shepherd, 2009), which involves land management practices, accelerated weathering, albedo modification, etc., and also technologies to reduce carbon levels from the atmosphere (Keith, 2009; Lackner et al., 2012).

At present, one can achieve negative emissions by co-firing excess biomass in a coal-fired power plant and capturing the CO<sub>2</sub> together with the CO<sub>2</sub> from the coal. This Bio-energy CCS (BECCS) scheme has an environmental advantage that the flue gas of the biomass is cleaned using the existing infrastructure of the coal-fired power plant (Gough and Upham, 2011). However, there is not enough disposable biomass that this can be applied on a large-scale.

## UTILIZATION

In the context of flue gases, CO<sub>2</sub> is seen as a waste product. However, there are many applications, where CO<sub>2</sub> is utilized or considered as a valuable commodity.

### ENHANCED OIL RECOVERY

At present, CO<sub>2</sub> is most valuable for enhanced oil recovery (EOR). In 2008, in the United States, about 80% of CO<sub>2</sub> for EOR is obtained from natural resources and the rest is from anthropogenic sources such as coal gasification or gas processing (Advanced Resources International (ARI), 2010). The permanent storage of CO<sub>2</sub> in depleted oil fields is definitely one of the attractive carbon storage options. The fact that there is a market for CO<sub>2</sub> is an important incentive to develop more efficient carbon capture technologies, such that anthropogenic CO<sub>2</sub> can compete with CO<sub>2</sub> from natural reservoirs.

There are two practical issues. Firstly, the total amount of CO<sub>2</sub> that can be used for EOR is much less than the total emissions of CO<sub>2</sub>, which implies that EOR can only be a partial solution. For example, CO<sub>2</sub> used in EOR operations was limited to only 60 million tons (Advanced Resources International (ARI), 2010). Secondly, one

may wonder whether EOR gives a net CO<sub>2</sub> reduction. The argument is that because of EOR we produce more oil, and hence, further increase anthropogenic CO<sub>2</sub> emissions. However, a better argument is to compare one barrel of oil produced with EOR compared to one barrel of oil that is produced without EOR. In this comparison, one barrel of oil with EOR gives a lower CO<sub>2</sub> emission as a fraction of the CO<sub>2</sub> used to recover the oil stays in the reservoir. However, here we did assume that the other barrel of oil does remain in the reservoir and that EOR does not increase the demand for oil.

### CO<sub>2</sub> TO CHEMICALS

If we look at the current chemical industry, about 7% of all oil is used as feedstock for carbon in products ranging from plastics to soaps. Replacing oil by a renewable feedstock is an important long-term challenge of the chemical industry. The viability of using CO<sub>2</sub> as a chemical feedstock is considerably improved if the price of carbon is sufficiently high such that CO<sub>2</sub> can replace oil. While the carbon-free energy sources required for CO<sub>2</sub>-to-chemical technologies (e.g., solar and wind) are still expensive for such CO<sub>2</sub> utilization schemes, the research and development of those CO<sub>2</sub> conversion pathways should be developed now to prepare for our rapidly changing future.

### CO<sub>2</sub> TO FUELS

The challenge with upgrading CO<sub>2</sub> to a fuel is that it requires energy. As it does not make any sense to use fossil fuels for this process, we assume that we will use renewable energy. The first argument is, if we have renewable energies we should primarily use this for generating electricity. However, this leaves us with two problems: storage of energy and transportation fuels.

An important advantage of fossil fuels is their high energy density. Renewable energies such as wind or solar require large-scale energy storage to ensure that electricity can be produced at time in which there is no wind or sun. Sometimes, this energy storage can be as simple as pumping water, but not all countries have this option. For example, Denmark has an excess of wind energy during the winter, but too little during summer. The idea is to use methanol to store the excess energy in the winter and use a conventional power plant with carbon

capture in the summer. In this cycle, efficient conversion of CO<sub>2</sub> into a fossil fuel like methanol is an essential step.

### INCORPORATING CO<sub>2</sub> INTO CONSTRUCTION AND BUILDING MATERIALS

The cement industry produces about 7% of CO<sub>2</sub> emissions and is the second largest emitter of CO<sub>2</sub> after coal-fired power plants (International Energy Agency Greenhouse Gas R&D Programme (IEA-GHG), 2002). Replacing 10% of building materials with carbonate minerals is expected to reduce CO<sub>2</sub> emissions by 1.6 Gt/year, which is about 5% of the global CO<sub>2</sub> emissions as of 2011 (Sridhar and Hill, 2011). However, it is important to determine the correct composition of carbonate minerals to be included in the concrete matrix to reduce issues related to mechanical strength of the materials.

### STORAGE

Given the enormous amounts of CO<sub>2</sub> we are emitting, it is difficult to imagine any form of carbon storage other than injecting into geological formations. Appropriate geological formations such as deep saline aquifers, depleted oil and gas fields, unmineable coal seams, and silicate formations (e.g., basalt) can accommodate up to 11,000 Gt CO<sub>2</sub> (Dooley et al., 2006), which is much greater compared to the annual CO<sub>2</sub> emissions, which are to the order of 30 Gt of CO<sub>2</sub>/year. In addition, from our experience with EOR we know how to transport and inject CO<sub>2</sub> in geological formations. The challenge is, however, the scale. At present, only about 50 Mt CO<sub>2</sub> has been stored today and 13 Mt CO<sub>2</sub>/year is expected by 2016 given plans in place for additional projects (Levina et al., 2013).

However, the scientific challenge is to ensure that the CO<sub>2</sub> remains safely in these storage sites for thousands of years. Development of technologies for monitoring, verification, and assessment (MVA) to ensure that the CO<sub>2</sub> remains trapped underground is essential. While the process of injecting CO<sub>2</sub> is well understood, the cost of monitoring the fate of injected CO<sub>2</sub> over many years may be too prohibitive unless, the cost of deploying MVA technologies is substantially reduced. In addition, key questions related to long-term safety such as induced seismicity and the potential for forming fractures need to be addressed,

which constitute an important aspect of risk assessments of geologic storage.

Ideally, mineralizing CO<sub>2</sub> into the form of a carbonate (e.g., limestone, magnesite) will reduce the amount of mobile CO<sub>2</sub> that needs to be monitored. However, the kinetics of this natural process may be on the order of geological timescales. An active area of research is to enhance this mineralization process (Gadikota et al., 2014).

## OUTLOOK

Given all the uncertainties, we argue that the first, and arguably, the most important challenge is to ensure that research is carried out on all aspects of CCUS. To reduce CO<sub>2</sub> levels in the atmosphere, land management and bio-energy with CCS (BECCS) may be low-hanging fruit. We can further envision capture technologies that are highly optimized for point sources ranging from coal-fired and gas-fired power plants to cement plants as well as direct air capture. The parasitic energy consumption should be lowered and the long-term recyclability of the CO<sub>2</sub> capture medium should be achieved. Cutting-edge research in carbon storage and utilization (i) will improve our understanding of the long-term effects of large-scale injection of CO<sub>2</sub> in geological formation, (ii) will enable us to develop alternatives for geological storage such as carbon mineralization, and (iii) will even allow for the development of the innovative chemistry to convert CO<sub>2</sub> into synthetic fuels and chemicals.

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