

Available online at [www.sciencedirect.com](http://www.sciencedirect.com)**SciVerse ScienceDirect**

Energy Procedia 12 (2011) 450 – 455

Energy

**Procedia**

ICSGCE 2011: 27–30 September 2011, Chengdu, China

## Effect of Oxygen Partial Pressure on the Optical Property of Amorphous Titanium Oxide Thin Films

Yongfeng Ju<sup>\*</sup>, Lin Li, Zhiming Wu, Yadong Jiang*State Key Laboratory of Electronic Thin Films and Integrated Devices, School of Optoelectronic Information, University of Electronic Science and Technology of China, Chengdu, China*

---

### Abstract

Amorphous thin films of titanium oxide (TiO<sub>x</sub>) have been deposited on K9 glass substrates by dc reactive magnetron sputtering from a pure titanium target at different oxygen partial pressure. The structure and composition of these films have been characterized by x-ray diffraction (XRD), and x-ray photoelectron spectroscopy (XPS), respectively. The effects of the oxygen partial pressure on the composition and optical properties of amorphous TiO<sub>x</sub> thin films have been mainly investigated. Analyses of transmittance, reflectance, optical band gap of the films demonstrate that the oxygen partial pressure during the sputtering plays a very important role in their optical properties.

© 2011 Published by Elsevier Ltd. Open access under [CC BY-NC-ND license](https://creativecommons.org/licenses/by-nc-nd/4.0/).

Selection and/or peer-review under responsibility of University of Electronic Science and Technology of China (UESTC).

*Keywords: amorphous titanium oxide; magnetron sputtering; optical properties*

---

### 1. Introduction

Titanium dioxide of non-stoichiometric (TiO<sub>2-δ</sub>) composition containing an excess of titanium is an n type semiconductor [1]. TiO<sub>x</sub> has a wide band-gap and in thin film form presents very attractive features due to its high refractive index and transparency in a wide spectral range (350–1100 nm). Besides, this material shows good stability in adverse environments, all these properties are of great interest in various technological applications such as photovoltaics [2], electrochromics [3], chemical sensors [4], photocatalysis [5]–[6], optoelectronic devices [7], dielectric material of ultrathin capacitors [8] and self-cleaning windows [9].

In particular, amorphous TiO<sub>x</sub> nanoparticles are only under intensive investigations for recent years by both of experiment [10] and computer simulation [11]. Though several methods have been developed to prepare amorphous TiO<sub>x</sub> nanoparticles, such as aqueous peroxotitanate solution method, [12],

---

<sup>\*</sup> Corresponding author. Tel.: +86-28-83206123.

E-mail address: [yfju\\_cd@yahoo.com.cn](mailto:yfju_cd@yahoo.com.cn).

microemulsion method [13], and supercritical carbon dioxide method [14], magnetron sputtering [15], very little is known about the optical properties of amorphous phase of  $\text{TiO}_x$  prepared by dc magnetron sputtering, which technique is of special interest because it is an industrial process applicable to large-area deposition and high quality films can be achieved even at low substrate temperatures. When the substrate temperature during deposition is low, amorphous  $\text{TiO}_x$  films are often observed [16].

In this paper, we report a study of effect of the oxygen partial pressure on some optical properties and the relation between the composition and the optical property of titanium oxide thin films prepared by dc reactive magnetron sputtering using an argon/oxygen mixture. The coatings have been deposited on K9 glass substrates at room temperature in order to obtain amorphous phase of  $\text{TiO}_x$  thin films.

## 2. Experimental Details

$\text{TiO}_x$  thin films were deposited on K9 glass at room temperature with a conventional dc magnetron sputtering system consisting of a stainless-steel vacuum chamber ( $150 \text{ dm}^3$ ) evacuated with a turbomolecular pump backed by a mechanical pump giving an ultimate vacuum of order of  $1 \times 10^{-4}$  Pa. The magnetron source, running at a dc power of 150 W, was placed at a distance of 60 mm from the substrate and tilted  $30^\circ$  away from the substrate normal. A resistive heater was mounted above the substrate so that the temperature, measured by a chromel-alumel thermocouple, could be varied from room temperature up to  $800^\circ\text{C}$ . The water cooled Ti metal target (purity 99.995%, GRIKIN) of 8 cm was used for depositing the films. The mixed reactive sputtering gas of Ar (99.999%) and  $\text{O}_2$  (99.995%) were controlled by two gas mass flux controllers, respectively. During sputtering, the total pressure was kept at 1.5 Pa, and the flux of Ar was 40 sccm (standard cubic centimeter per minute) while the flux of  $\text{O}_2$  varied from 2 to 6 sccm. Prior to all the depositions, the substrates were ultrasonically cleaned with acetone, ethanol, and de-ionized water by an ultrasonic cleaning machine for 10 minutes and then were blown dry with nitrogen.

The phases present in the  $\text{TiO}_x$  thin films were analyzed by an X-ray diffractometer (X'Pert Pro MPD, PANalytical B.V.) using  $\text{Cu K}\alpha$  radiation ( $\lambda = 0.15406 \text{ nm}$ ) and operating at an accelerating voltage of 40 kV and an emission current of 35 mA. Data were acquired over the range from  $20$  to  $70^\circ$  at a sampling width of  $0.03^\circ$ . The XRD method was used to study the change of crystalline structure. Thickness of the as-deposited films was measured by profilometer (XP-2, Ambios Technology Inc.). The stoichiometry of titanium oxides was determined by x-ray photoelectron (XPS) spectra using x-ray photoelectron spectrometer (XSAM 800, Kratos) at room temperature with monochromatic  $\text{Al K}\alpha$  ( $h\nu = 1486 \text{ eV}$ ). The vacuum of the instrument chamber was  $1 \times 10^{-7}$  Pa. The binding energy was calibrated with reference to C 1s peak (285.0 eV). The optical transmission and reflection spectra of  $\text{TiO}_x$  thin films were recorded at room temperature by a double-beam spectrophotometer (UV-2550, Shimadzu Co.). The wavelength used in the experiment ranged from 300 to 800 nm.

## 3. Results and discussion

### 3.1. X-ray Diffraction (XRD)

Fig. 1 shows the XRD pattern of as-deposited  $\text{TiO}_x$  films prepared at different oxygen partial pressure. As showed in Fig. 1, no sharp diffraction peaks of any crystalline phase appear in the deposited films, this indicates that all the films deposited of different oxygen partial pressure at room temperature are amorphous in structure. This result is in good agreement with the literature [17].

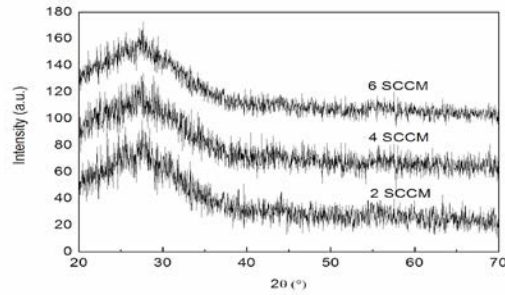


Fig. 1. XRD pattern of  $\text{TiO}_x$  thin films deposited at different oxygen partial pressure.

### 3.2. X-ray Photoelectron Spectra (XPS) Measurements

Fig. 2 shows the XPS spectra of Ti 2p energy level of  $\text{TiO}_x$  films deposited at the different oxygen partial pressure of 2 sccm and 6 sccm, respectively. The C 1s peak at 285.0 eV is used to calibrate the spectra. The spectra in Fig. 2(a) can be resolved into four Gaussian components. The Ti 2p<sub>3/2</sub> peak at 457.5 eV and Ti 2p<sub>1/2</sub> peak at 463.0 eV are attributed to  $\text{Ti}^{3+}$  while the peak Ti 2p<sub>3/2</sub> at 459.1 eV and Ti 2p<sub>1/2</sub> at 464.8 eV are attributed to  $\text{Ti}^{4+}$  [18]. This indicates that titanium oxide ( $\text{TiO}_x$ ) but not titanium dioxide ( $\text{TiO}_2$ ) films can be obtained by dc magnetron sputtering at the lower oxygen partial pressure. It is also clear that the Ti ions in sputtered  $\text{TiO}_x$  films are prominent in charge state of  $\text{Ti}^{4+}$ . The spectra in Fig. 2(b) can be only resolved into two Gaussian components. The Ti 2p<sub>3/2</sub> peak at 458.7 eV and the peak Ti 2p<sub>1/2</sub> at 464.3 eV are attributed to  $\text{Ti}^{4+}$ , and no peaks attributed to  $\text{Ti}^{3+}$  ions appear. Comparing the two XPS spectra of deposited  $\text{TiO}_x$  films, it can be concluded that with the increasing of oxygen partial pressure, the oxygen content in  $\text{TiO}_x$  increases, too. In other words, using different oxygen partial pressure in reactive magnetron sputtering, different titanium oxide films could be obtained, as we all known, this is one of the merits for reactive sputtering. Therefore, we could control the stoichiometry of metal oxide films by this method, which is very important in the metal oxide semiconductor's electrical property [1].

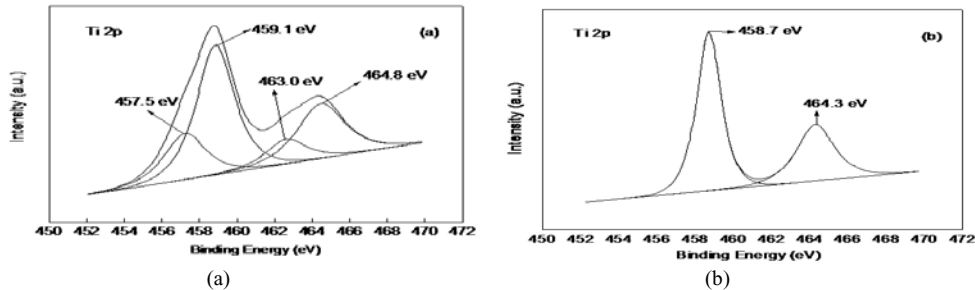


Fig. 2. The XPS spectra of Ti 2p energy level of  $\text{TiO}_x$  thin films deposited at different oxygen partial pressure: (a) 2 sccm, (b) 6 sccm.

### 3.3. Optical Properties

Fig. 3 and Fig. 4 show the transmission and reflectance spectra of as-deposited  $\text{TiO}_x$  thin films at different oxygen partial pressure. It can be seen that the transmittance of the  $\text{TiO}_x$  thin films increases as the oxygen partial pressure increases. When the oxygen partial pressure increases, the film structure changes from dense into porous, in this case, the film has more voids and part of the light can pass

through the film without scattering and results in an increase of the transmittance. As shown in Fig. 4, the position of the reflectivity minimum shifts to longer wavelength, i.e., lower energy with the the increasing of oxygen partial pressure. This result is in good agreement with XPS analysis and the literature [19], which provides a quick way of determining the composition of the  $\text{TiO}_x$  films.

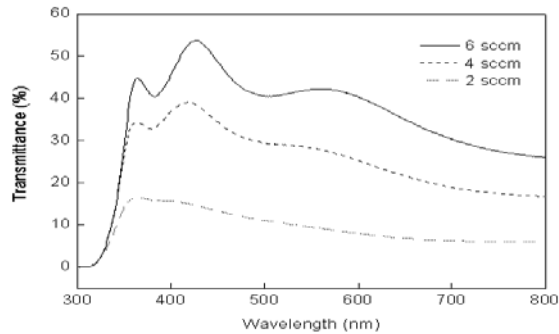


Fig. 3. Optical transmittance as a function of wavelength of  $\text{TiO}_x$  thin films deposited at different oxygen partial pressure.

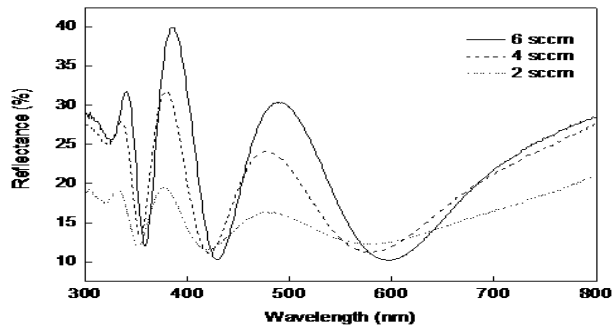


Fig. 4. Optical reflectance as a function of wavelength of  $\text{TiO}_x$  thin films deposited at different oxygen partial pressure.

Since titanium dioxide is a semiconductor with a large band gap, the optical band gap  $E_g$  can be determined from absorption coefficient  $\alpha$ . The absorption coefficient  $\alpha$ , which depends on the wavelength  $\lambda$ , can be obtained by using the relation [20].

$$T = (1-R)^{2\exp(-\alpha d)} \quad (1)$$

where  $T$  is the transmittance,  $R$  is the reflectance,  $\alpha$  is the absorption coefficient and  $d$  is the film thickness which was measured by the profilometer. As the increasing of oxygen partial pressure, the thickness of the  $\text{TiO}_x$  thin films decrease obviously, i.e., 300 nm (2 SCCM), 280 nm (4 SCCM), 255 nm (6 SCCM), respectively. The drop in deposition rate is due to oxidation of the target, which is a typical characteristic of dc reactive sputtering.

Equation (1) is valid close to the optical band gap of the material under the condition  $\exp(2\alpha d) \gg R^2$ . From Equation (1), the absorption coefficient  $\alpha$  could be deduced by:

$$\alpha = -d^{-1} \ln [T / (1-R)^2] \quad (2)$$

Above the threshold of fundamental absorption, the dependence of  $\alpha$  on incident light energy is given by the expression [21]:

$$\alpha h\nu = \alpha_0(E-E_g)^m \quad (3)$$

where  $E=h\nu$  is the photon energy,  $E_g$  is the optical band gap and  $\alpha_0$  is a constant which does not depend on  $E$ . Since  $\text{TiO}_x$  is a semiconductor with a direct band gap but dipole-forbidden transitions, the value of  $m$  can be used as  $3/2$  [22].

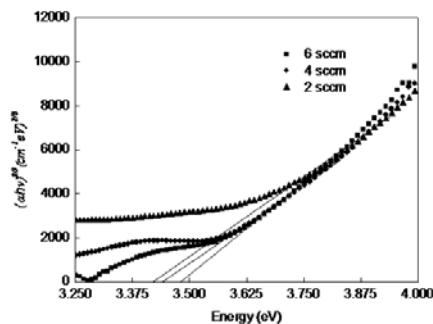


Fig. 5. Optical band gaps for as-deposited  $\text{TiO}_x$  thin films at different oxygen partial pressure.

From the Fig. 5, it is observed that optical band gap for the as-grown  $\text{TiO}_x$  thin film at room temperature decreases from 3.45 eV to 3.42 eV, which is not obvious with the increase of oxygen partial pressure in our experiments. However, these results have shown clear optical blueshift of the absorption edges comparing to bulk  $\text{TiO}_2$  (Anatase: 3.2 eV, Rutile: 3.0 eV) which can be explained by the quantum-size (Q-size) effect, i.e., the small size of titanium oxide nanoparticles raises the conduction band and lowers the valance band, finally cause a blueshift of the band gap [23]-[24].

#### 4. Conclusions

Amorphous titanium oxide ( $\text{TiO}_x$ ) thin films have been prepared by dc reactive magnetron sputtering at different oxygen partial pressure. The composition and optical properties of the sputtered films such as transmittance, reflectance and optical band gap of the sputtered films have been investigated, i.e., with the increasing of oxygen partial pressure, the oxygen content of the  $\text{TiO}_x$  films increases, which make the transmittance of the  $\text{TiO}_x$  thin films increase and the position of reflectivity minimum has a red shift. Because of quantum-size effect, the optical band gap of deposited  $\text{TiO}_x$  films is found to have a blueshift comparing to bulk  $\text{TiO}_x$ , and with the increasing oxygen partial pressure, it decreases from 3.45 eV to 3.42 eV.

#### Acknowledgment

This work is financially supported by National Science Foundation of China of Grant No. 60736005.

#### References

- [1] R. G. Breckenridge and W. R. Hosler, "Electrical Properties of Titanium Dioxide Semiconductors," *Phys. Rev.*, vol. 91, 1953, pp. 793-802.
- [2] B. O'Regan and M. Gratzel, "A low-cost, high-efficiency solar cell based on dye-sensitized colloidal  $\text{TiO}_2$  films," *Nature*, vol. 353, 1991, pp. 737-740.
- [3] U. Bach, D. Corr, D. Lupo, F. Pichot and M. Ryan, "Nanomaterials-Based Electrochromics for Paper-Quality Displays," *Adv. Mater.*, vol. 14, 2002, pp. 845-848.

- [4] L. D. Birkefeld, A. M. Azadand and S. A. Akbar, "Carbon Monoxide and Hydrogen Detection by Anatase Modification of Titanium Dioxide," *J. Am. Ceram. Soc.*, vol. 75, 1992, pp. 2964-2968.
- [5] A. Fujishima and K. Honda, "Electrochemical Photolysis of Water at a Semiconductor Electrode," *Nature*, vol. 238, 1972, pp. 37-38.
- [6] D. C. Hurum, A. G. Agrios, K. A. Gray, T. Rajh and M. C. Thurnauer, "Explaining the Enhanced Photocatalytic Activity of Degussa P25 Mixed-Phase TiO<sub>2</sub> Using EPR," *J. Phys. Chem. B*, 2003, vol. 107, pp. 4545-4549.
- [7] P. Ravirajan, S. A. Haque, J. R. Durrant, D. D. C. Bradley and J. Nelson, "The Effect of Polymer Optoelectronic Properties on the Performance of Multilayer Hybrid Polymer/TiO<sub>2</sub> Solar Cells," *Adv. Funct. Mater.*, 2005, Vol. 15, pp. 609-618.
- [8] T. Brezesinski, J. Wang, J. Polleux, B. Dunn and S. H. Tolbert, "Templated Nanocrystal-Based Porous TiO<sub>2</sub> Films for Next-Generation Electrochemical Capacitors," *J. Am. Chem. Soc.*, 2009, vol. 131, pp. 1802-1809.
- [9] O. Bikondoa, C. L. Pang, R. Ithnin, C. R. Muryn, H. Onishi, and G. Thornton, "Direct visualization of defect-mediated dissociation of water on TiO<sub>2</sub>(110)," *Nat. Mater.*, 2006, vol. 5, pp. 189-192.
- [10] V. Luca, S. Djajanti and R. F. Howe, "Structural and Electronic Properties of Sol-Gel Titanium Oxides Studied by X-ray Absorption Spectroscopy," *J. Phys. Chem. B*, 1998, vol. 102, pp. 10650-10657.
- [11] K. T. Lim, H. S. Hwang, W. Ryoo and K. P. Johnston, "Synthesis of TiO<sub>2</sub> Nanoparticles Utilizing Hydrated Reverse Micelles in CO<sub>2</sub>," *Langmuir*, 2004, vol. 20, pp. 2466-2471.
- [12] Y. Gao, Y. Masuda, Z. Peng, T. Yonezawa and K. Koumoto, "Room temperature deposition of a TiO<sub>2</sub> thin film from aqueous peroxotitanate solution," *J. Mater. Chem.*, 2003, vol. 13, pp. 608-613.
- [13] M. S. Lee, S. S. Park, G. Lee, C. Ju and S. Hong, "Synthesis of TiO<sub>2</sub> particles by reverse microemulsion method using nonionic surfactants with different hydrophilic and hydrophobic group and their photocatalytic activity," 2005, *Catal. Today*, vol. 101, pp. 283-290.
- [14] C. Wu, J. Huang, Y. Wen, S. Wen, Y. Shen and M. Yeh, "Preparation of TiO<sub>2</sub> nanoparticles by supercritical carbon dioxide," *Mater. Lett.*, 2008, vol. 62, pp. 1923-1926.
- [15] N. Martin, C. Rousselot, C. Savall and F. Palmino, "Characterizations of titanium oxide films prepared by radio frequency magnetron sputtering," *Thin Solid Films*, 1996, vol. 287, pp. 154-163.
- [16] P. Lobl, M. Huppertz and D. Mergel, "Nucleation and growth in TiO<sub>2</sub> films prepared by sputtering and evaporation," *Thin Solid Films*, 1994, vol. 251, pp. 72-79.
- [17] M. Radecka, K. Z. Akrczewska, H. Cztemastek, T. Stapinski and S. Debrus, "The influence of thermal annealing on the structural, electrical and optical properties of TiO<sub>2-x</sub> thin films," *Appl. Surf. Sci.*, 1993, vol. 65-66, pp. 227-234.
- [18] A. Manivannan, G. Glaspell and M. S. Seehra, "Controlled transformation of paramagnetism to room-temperature ferromagnetism in cobalt-doped titanium dioxide," *J. Appl. Phys.*, 2003 vol. 94, pp. 6994-6996.
- [19] O. Banakh, P. E. Schmid, R. Sanjinès and F. Lévy, "Electrical and optical properties of TiO<sub>x</sub> thin films deposited by reactive magnetron sputtering," *Surf. Coat. Technol.*, 2002, vol. 151, pp. 272-275.
- [20] T. Asanuma, T. Matsutani, C. Liu, T. Mihara, and M. Kiuchi, "Structural and optical properties of titanium dioxide films deposited by reactive magnetron sputtering in pure oxygen plasma," *J. Appl. Phys.*, 2004, vol. 95, pp. 6011-6016.
- [21] J. Tauc, *Amorphous and Liquid Semiconductors*, London, Plenum, 1974, pp. 175-179.
- [22] N. Daude, C. Gout and G. Jouanin, "Electronic band structure of titanium dioxide," *Phys. Rev. B*, 1977, vol. 15, pp. 3229-3235.
- [23] M. Anpo, T. Shima, S. Kodama and Y. Kubokawa, "Photocatalytic hydrogenation of propyne with water on small-particle titania: size quantization effects and reaction intermediates," *J. Phys. Chem.*, 1987, vol. 91, pp. 4305-4310.
- [24] C. Kormann, D. W. Bahnemann and M. R. Hoffmann, "Preparation and characterization of quantum-size titanium dioxide," *J. Phys. Chem.*, 1988, vol. 92, pp. 5196-5201.