

## Carbon-nitrogen/graphene composite as metal-free electrocatalyst for the oxygen reduction reaction

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Sheet-like carbon-nitrogen (CN<sub>x</sub>)/graphene composites with a high content of nitrogen ( $x \leq 0.15$ ) was prepared by the carbonization of polypyrrole (PPy)/reduced-graphene-oxide (rGO) composite at 600–800°C. We used rGO instead of graphene oxide (GO) sheets as a template and a substrate to immobilize PPy since the PPy/GO composite agglomerates easily because of the dehydration of excess oxygen-containing groups on the GO sheets during the drying process. The dried PPy/rGO intermediate and its derived CN<sub>x</sub>/graphene products retain their high dispersion and loose-powder features. The as-prepared CN<sub>x</sub>/graphene composites have a total nitrogen content of about 10 at% and their nitrogen state is mainly of pyridinic and graphitic type. CN<sub>x</sub>/graphene composites exhibit excellent performance for the oxygen reduction reaction (ORR) in terms of electrocatalytic activity, stability and immunity towards methanol crossover and CO poisoning, suggesting their potential as metal-free electrocatalysts for the ORR.

**graphene, graphene oxide, carbon nitride, metal-free electrocatalyst, oxygen reduction reaction**

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Low temperature fuel cells are a high-efficiency and environment-friendly energy supply for future transport and portable applications [1,2]. However, despite the great efforts worldwide over the past few decades their wide application in daily life is still a challenge because of the scarce resources, high cost and durability issues for the current commercial Pt/C catalysts, especially for cathodes where the oxygen reduction reaction (ORR) is kinetically slow [3–5]. Therefore, research into non-precious metals or even metal-free catalysts with high ORR activity and durability is absolutely crucial for the development of fuel cells.

Currently, it is widely accepted that nitrogen-containing carbon including macrocycle molecules [6–8], polymers [4,9] and nanomaterials [10–19] are a type of important building block for the preparation of Pt-substituted catalysts.

Iron- or cobalt-doped carbon-nitrogen (CN<sub>x</sub>) compounds have comparable activity to Pt/C catalysts in acidic systems [4,7,9,20]. Polypyrrole modified carbon-supported cobalt hydroxide can act as cathode and anode catalysts in direct borohydride fuel cells [21,22]. Recent promising results involving metal-free PEDOT [5] and nitrogen-doped carbon nanotubes (NCNTs) [14–19] have been reported for alkaline systems giving excellent electrocatalytic activity, stability and immunity towards methanol crossover and CO poisoning. Particularly, NCNTs have attracted extensive interest because of their large surface area, good thermal and chemical stability as well as high electrical conductivity [14–19]. It has been revealed that increasing the nitrogen content and the number of defects in NCNT enhances its electrocatalytic activity toward the ORR [11,16]. In addition to one-dimensional CNTs, the emergence of graphene has opened up a new field in terms of research into two-dimensional (2D)

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metal-free electrocatalysts [23–29]. Graphene has many similar properties to CNTs and its 2D planar sheet is beneficial for electron transport and, therefore, it is an ideal electrode material [23,24]. Nitrogen doped graphene with a low nitrogen concentration of 2 at%–5 at% can be synthesized by chemical vapor deposition [25,30] or by the post-nitridization of graphene oxide [28,31]. To increase the nitrogen content, researchers have immobilized a nitrogen-containing polymer such as polyaniline [32] or polymelamine onto the graphene oxide (GO) surface to form polymer/GO composite [29], which can be converted to  $CN_x$ /graphene composite by carbonization to catalyze the ORR. The main challenge for the preparation of the intermediate of the polymer/GO-composite powder is to avoid agglomeration and caking after natural drying because of GO sheet dehydration [32]. Therefore, it is necessary to eliminate the excess oxygen-containing groups on the GO sheets by the reduction method while retaining the high dispersion of reduced GO (rGO) in solution. In this study, we prepared highly dispersed rGO solution by hydrothermal reduction and then polypyrrole (PPy) was added onto the rGO sheets to prepare the PPy/rGO composite. Finally, the fabricated  $CN_x$ /graphene composite with a nitrogen content of about 10 at% was prepared by simple carbonization. The obtained  $CN_x$ /graphene composite is shown to be a metal-free electrocatalyst for the ORR with good electrocatalytic activity, long-term operation stability and a high tolerance of the crossover effect, indicating their potential use in alkaline fuel cells.

## 1 Experimental

### 1.1 Synthesis of graphite oxide

Graphite oxide was synthesized by the modified Hummers method [33]. In short, 1 g natural graphite (325 mesh, Alfa Aesar) was added to 50 mL concentrated sulfuric acid while stirring in an ice-water bath. Then, 6 g potassium permanganate was added slowly and the mixture was stirred at 30°C for 1 h after which 80 mL distilled water was added and the mixture was further stirred for 30 min and the temperature was increased to 90°C. Finally, 200 mL distilled water and 6 mL  $H_2O_2$  (30%) were added dropwise and the solution changed color from dark brown to yellow. The generated solid graphite oxide was separated by centrifugation, washed and finally dried under vacuum.

### 1.2 Preparation of the PPy/rGO composite

Graphite oxide (40 mg) was dispersed in 40 mL deionized water by sonication to form a 1 mg/mL GO suspension, which was transferred into 50 mL autoclaves and heated at 150°C for 12 h. After the reduction, a black homogeneous rGO suspension was obtained. Pyrrole monomer (0.05 mL) was added into the rGO solution and magnetically stirred for 24 h at room temperature. Then, 50 mL of an aqueous

solution containing 189.5 mg ammonium peroxydisulfate was added as an oxidizing agent to initiate the polymerization for 24 h in an ice-bath. After the reaction, the solid sample was filtered and thoroughly washed with deionized water and ethanol alternately, and then it was dried naturally to form a PPy/rGO powder. For comparison, the PPy/GO composite was also prepared by the direct use of GO as a substrate without hydrothermal treatment.

### 1.3 Preparation of the $CN_x$ /graphene composite

PPy/rGO powder (0.1 g) was placed in a quartz tube furnace and heated under Ar from room temperature to 600 and 800°C at a rate of 1°C/min and kept there for 2 h. A black powder was obtained after the furnace was cooled to room temperature and this was denoted  $CN_x$ /G-600 or  $CN_x$ /G-800, respectively, according to the annealing temperature.

### 1.4 Characterization

The products were characterized by transmission electron microscopy (TEM, JEOL-JEM-1005 at 100 kV), scanning electron microscopy (SEM, Hitachi S-4800), high-resolution transmission electron microscopy (HRTEM, JEM2010 at 200 kV), Fourier transform infrared spectroscopy (FTIR, BRUKERVECTOR22) with KBr pressed pellets, Raman spectroscopy (Renishaw in Via Raman Microscope with an argon-ion laser at an excitation wavelength of 514 nm) and X-ray photoelectron spectroscopy (XPS, VG ESCALAB MKII).

Electrochemical measurements included rotating disk electrode (RDE) and rotating ring-disk electrode (RRDE) voltammetries were carried out at 25°C on a CHI 760 C workstation (CH Instruments). Ag/AgCl and a platinum wire served as reference and counter electrodes, respectively. The RDE and RRDE voltammetry experiments were performed on a MSR electrode rotator (Pine Instrument Co.). RDE voltammetry was conducted using a glassy carbon (GC) electrode (5 mm diameter) while for RRDE voltammetry we used a GC disk electrode (5 mm diameter) surrounded by a Pt ring (6.5 mm inside diameter). The GC electrode was modified as following: A suspension of catalyst with a concentration of  $1.0 \text{ mg mL}^{-1}$  was made by ultrasonically dispersing 1 mg catalyst in 1 mL water. Then, 100  $\mu\text{L}$  of the catalyst suspension was dropped onto a glassy carbon disk (GC) intermittently. 0.5  $\mu\text{L}$  Nafion (Dupont, 5 wt%) was then coated onto the surface of the thin film electrode and dried thoroughly in air for 12 h. ORR evaluation was performed in a  $0.1 \text{ mol L}^{-1}$  KOH electrolyte that was saturated and protected by  $O_2$ . The electrochemical performances of all the catalysts were examined determined in a  $0.1 \text{ mol/L}$  KOH solution.

Thin film electrodes for each catalyst were prepared for the methanol crossover and CO poisoning tests by the same procedure. Simply, 20  $\mu\text{L}$  of the catalyst suspension was

dropped onto a glassy carbon disk. After drying thoroughly, 0.5  $\mu\text{L}$  Nafion (Dupont, 5 wt%) was coated onto the surface of the thin film electrode. Methanol crossover and CO poisoning were studied at a polarizing potential of  $-0.3$  V through their  $i$ - $t$  chronoamperometric response in the oxygen saturated electrolyte by bubbling  $\text{O}_2$  into the electrolyte at a flowrate of 30 sccm (standard cubic centimeter per minute). During this process, 0.1 mL methanol or a 30 sccm (standard cubic centimeter per minute) flow of CO was introduced into the electrolyte over 400 s to examine methanol crossover and CO poisoning.

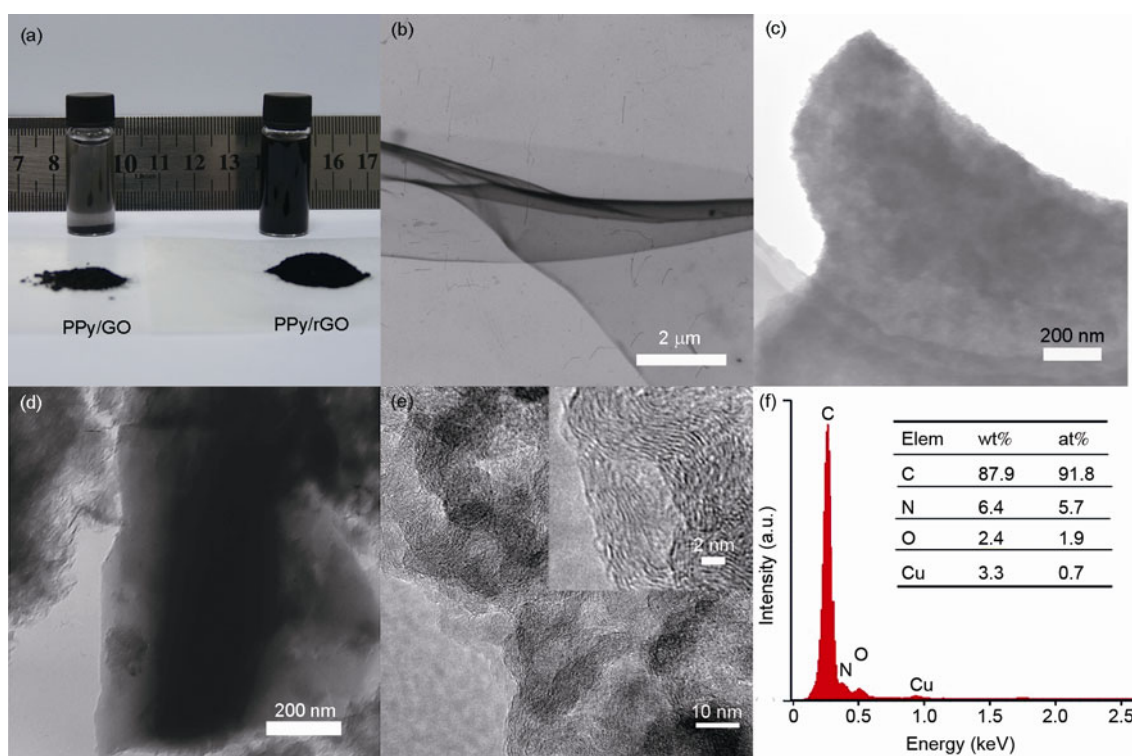
## 2 Results and discussion

Before using rGO as a substrate to immobilize PPy, we prepared a PPy/GO composite with a weight ratio of 1:1 and observed that the dried powder agglomerated into particles and could not be redispersed into water after 30 min of ultrasonication (Figure 1(a)). The composite "cake" was a result of dehydration among the unenclosed GO sheets during the drying process. To avoid this agglomeration, hydrothermally treated GO, i.e., rGO, instead of pristine GO was used as a matrix to support the polymer. According to our previous study [34], the C/O atomic ratio for GO is  $\sim 3.4$  while that for rGO is  $\sim 7.6$ . Reduced GO sheets with less

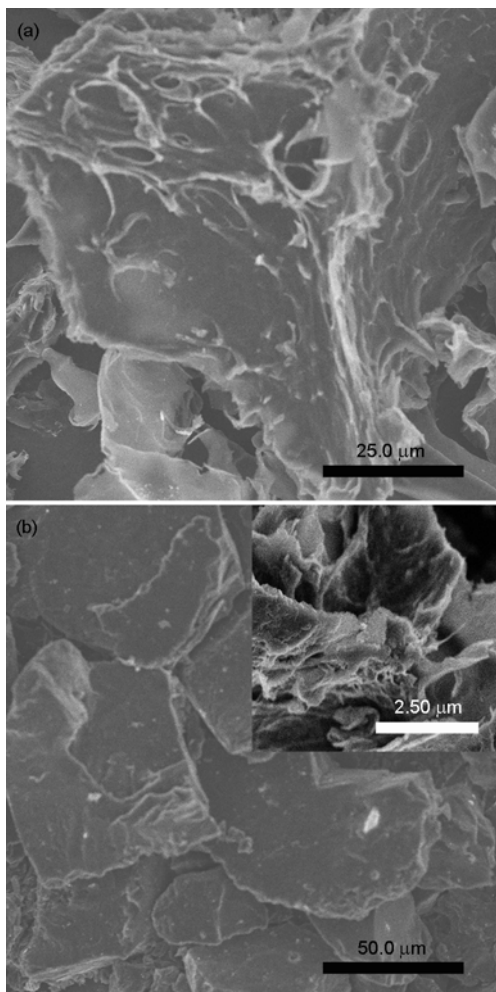
oxygen containing groups can also maintain high dispersion in solution in the form of free-standing flakes, as shown in Figure 1(b). Obviously, the dry PPy/rGO composite is much looser than PPy/GO and the former can easily form a homogeneous dispersion after several minutes of ultrasonication (Figure 1(a)). As shown by a typical TEM image of PPy/rGO (Figure 1(c)), the PPy polymer attaches tightly to the rGO sheet and forms a thicker plate than the pristine rGO sheet. After calcination, the samples retained their plate-like morphology. The HRTEM image of the calcinated sample shows the microstructure at the edge where the stacking structure is distinguishable and the discontinuous graphitic layers indicate the formation of graphite fragments. The EDS spectrum reveals N doping with an uncalibrated content of 5.7 at%, which indicates the formation of the  $\text{CN}_x/\text{graphene}$  composite.

SEM images of rGO and  $\text{CN}_x/\text{G-800}$  are shown in Figure 2. The image for rGO shows that the thin and wrinkled sheets are stacked tightly together. Compared with rGO, the  $\text{CN}_x/\text{G-800}$  sheets look much thicker and they are also stacked as composite flakes, as shown in the insert in Figure 2(b). This slight aggregation of the  $\text{CN}_x/\text{graphene}$  composite does not affect their dispersion in solution and a homogeneous suspension with high stability can be obtained for the preparation of a thin film electrode.

The functional groups on the surface of the rGO, PPy/

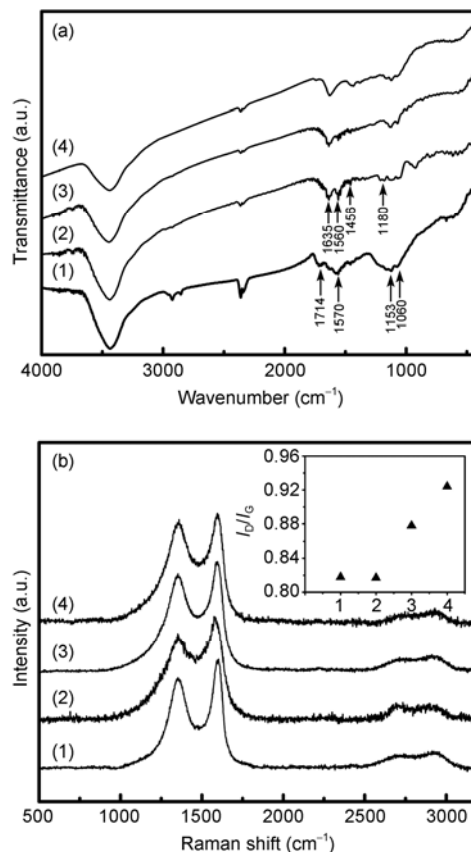


**Figure 1** (a) Photographs of the PPy/GO and PPy/rGO composite powders and their solutions. PPy/GO was not well dispersed in the solution after 30 min of ultrasonication while PPy/rGO tends to form a homogeneous dispersion after several minutes of ultrasonication; typical TEM images of rGO sheets (b), PPy/rGO (c) and  $\text{CN}_x/\text{G-800}$  (d); (e) HRTEM image of  $\text{CN}_x/\text{G-800}$ . The insert is an enlarged image; (f) EDS spectrum and the derived elemental composition of  $\text{CN}_x/\text{G-800}$ .



**Figure 2** SEM images of rGO (a) and  $\text{CN}_x/\text{G-800}$  (b). The insert in (b) is a local enlargement.

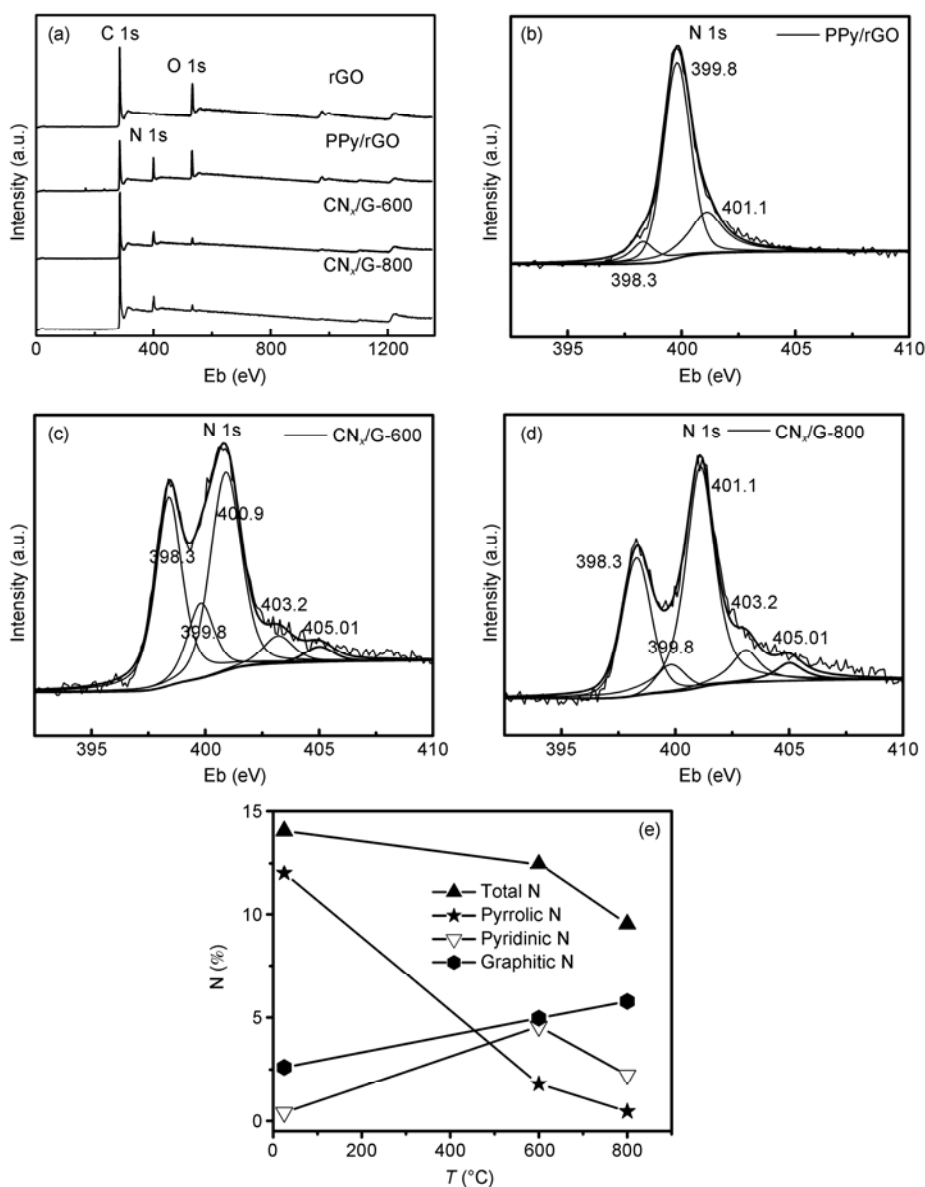
rGO and  $\text{CN}_x/\text{graphene}$  composites were analyzed by FTIR (Figure 3(a)). The spectrum of rGO gives typical peaks at  $1060\text{ cm}^{-1}$  for the C–O stretching vibration and  $1714\text{ cm}^{-1}$  for C=O [35]. The pyrrole ring vibration at  $1456$  and  $1560\text{ cm}^{-1}$  accompanied by the =C–H in-plane deformation at  $1180$  and  $1336\text{ cm}^{-1}$  appear in the FTIR spectrum of the PPy/rGO composite [36]. After annealing at  $600$  and  $800^\circ\text{C}$ , the characteristic peaks for PPy diminish while some peaks between  $1100$ – $1700\text{ cm}^{-1}$  come from C–N and C=N stretching vibrations [37]. Figure 3(b) shows Raman spectra of rGO, PPy/rGO and their derived carbon nitride composites. All the samples have spectra with two peaks centered at  $\sim 1353\text{ cm}^{-1}$  (D-band) and  $\sim 1594\text{ cm}^{-1}$  (G-band) [38]. The G-band is assigned to one of the two  $E_{2g}$  modes corresponding to stretching vibrations in the basal-plane ( $sp^2$  domains) of single crystal graphite or graphene. The D-band is usually attributed to disorders and imperfections in carbon crystallites. The intensity ratio of the D-band to G-band ( $I_D/I_G$ ) in the Raman spectrum should approximately correspond to the extent of disorder in the graphitic carbon [26,38].



**Figure 3** FTIR (a) and Raman spectra (b) of rGO (1), PPy/rGO (2),  $\text{CN}_x/\text{G-600}$  (3) and  $\text{CN}_x/\text{G-800}$  (4). The insert in (b) is the  $I_D/I_G$  ratio profile for each sample.

The  $I_D/I_G$  ratio of  $\text{CN}_x/\text{graphene}$  is higher than that of rGO indicating that  $\text{CN}_x/\text{graphene}$  has a higher distortion.

The chemical state evolution of nitrogen before and after calcination was analyzed by XPS as shown in Figure 4. From the XPS survey scan in Figure 4(a), the spectrum of rGO does not have a N signal as expected. The relative intensities of the N1s and O1s peaks for PPy/rGO,  $\text{CN}_x/\text{G-600}$  and  $\text{CN}_x/\text{G-800}$  decrease sequentially because of the higher annealing temperature. The N1s spectrum for PPy/rGO,  $\text{CN}_x/\text{G-600}$  and  $\text{CN}_x/\text{G-800}$  can be fitted to individual component peaks that represent pyrrolic N at a binding energy ( $E_b$ ) of  $399.8\text{ eV}$ , pyridinic N at a  $E_b$  of  $398.3\text{ eV}$  and graphitic N at a  $E_b$  between  $400.9$ – $401.1\text{ eV}$  [26,27,31]. Additionally, the N1s spectra of the carbonized sample (Figure 4(c) and (d)) also contain two other peaks at  $E_b$ s of  $403.2$  and  $405.0\text{ eV}$ , corresponding to N-oxide and intercalated  $\text{N}_2$ , respectively [39]. The total N content obtained from XPS as well as the pyrrolic, pyridinic and graphitic nitrogen content from the relative peak intensities depends on the annealing temperature, as plotted in Figure 4(e). For the as-prepared PPy/rGO composite at room temperature, approximately 90% of the nitrogen exists as pyrrolic N and the total N content is about 14 at%. With an increase in the annealing temperature, the total N and pyrrolic N decreases while the

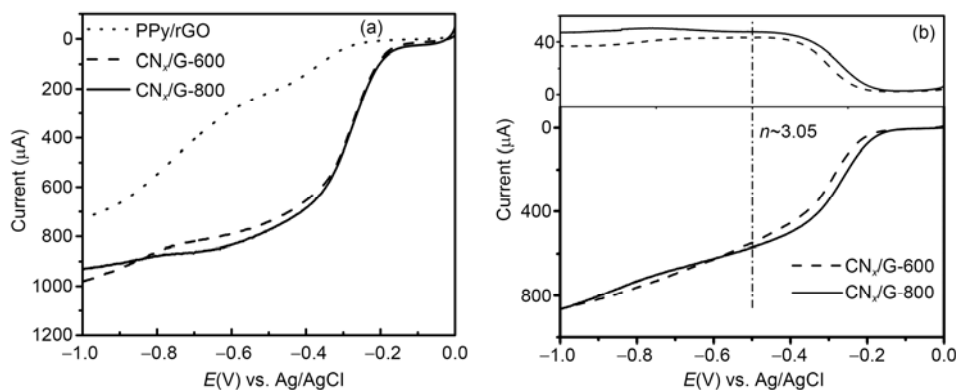


**Figure 4** XPS spectra of rGO, PPy/rGO, CN<sub>x</sub>/G-600 and CN<sub>x</sub>/G-800. Survey scan (a); N1s narrow scan of PPy/rGO (b), CN<sub>x</sub>/G-600 (c) and CN<sub>x</sub>/G-800 (d); (e) the total N content, the derived pyrrolic N, the pyridinic N and the graphitic N vs. the annealing temperature.

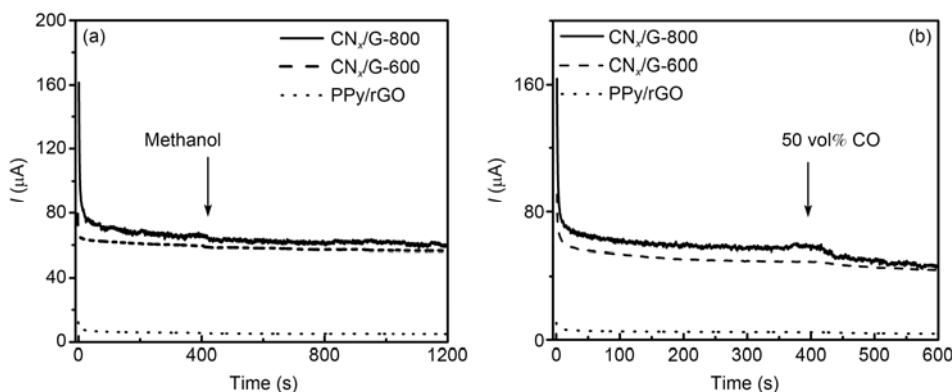
graphitic N increases. The pyridinic N reaches its highest content at 600°C. The total N content for CN<sub>x</sub>/G-600 and CN<sub>x</sub>/G-800 are 12.5 at% and 9.8 at%, respectively.

The electrocatalytic activities of the CN<sub>x</sub>/graphene composites were determined using RDE and RRDE voltammeteries (Figure 5). For RDE (Figure 5(a)) the onset potential of oxygen reduction for the CN<sub>x</sub>/graphene composites are about  $-1.9$  V, which is much higher than that of the PPy/rGO composite ( $\sim -0.3$  V). Moreover, at a given applied potential, the ORR current of the CN<sub>x</sub>/graphene composite is obviously stronger than that of the PPy/rGO composite. These results indicate that pyridinic and graphitic nitrogen rather than pyrrolic nitrogen contributes significantly to the ORR catalytic activity [27,28]. The activity of the carbon-nitrogen metal-free materials is synergistically deter-

mined by nitrogen content, the chemical state of nitrogen and the electrical conductivity. Although the CN<sub>x</sub>/G-600 catalyst has a higher amount of total and pyridinic nitrogen than the CN<sub>x</sub>/G-800 catalyst, the latter has a smaller resistance, i.e., higher conductivity than the former according to the electrochemical impedance measurements (Figure S2). Therefore, CN<sub>x</sub>/G-600 and CN<sub>x</sub>/G-800 have similar activity. The transferred electron number ( $n$ ) per oxygen molecule involved in the ORR was calculated using a RRDE measurement (Figure 5(b)) and  $n = 4I_{\text{disk}} / (I_{\text{disk}} + I_{\text{ring}}/N)$  [15], where  $I_{\text{disk}}$  and  $I_{\text{ring}}$  are the faradaic disk current and the faradaic ring current while  $N = 0.26$  is the collection efficiency. Here  $n$  was calculated at a polarized potential of  $-0.5$  V, which is 3.03 for CN<sub>x</sub>/G-600 and 3.07 for CN<sub>x</sub>/G-800, and this is higher than that of the CN<sub>x</sub>/graphene composite prepared by



**Figure 5** RED (a) and RRED (b) voltammeteries measured in an  $O_2$ -saturated 0.1 mol/L KOH electrolyte with a rotation speed of 2000 r/min and a scan rate of  $10 \text{ mV s}^{-1}$ . In the RRDE tests, the Pt ring electrode was polarized at 0.5 V.



**Figure 6** The  $i$ - $t$  chronoamperometric responses for PPY/rGO,  $CN_x$ /G-600 and  $CN_x$ /G-800. (a) Methanol crossover tests after the introduction of 0.1 mL methanol to the electrolyte at 400 s. (b) CO poisoning tests by introducing additional CO at the same  $O_2$  flow rate into the electrolyte at 400 s.

the polymerization of melamine ( $n = 2.5$ – $2.7$ ) [29]. This may be caused by the good graphitization that results from the higher calcination temperature used.

Another prominent advantage of the metal-free catalyst for the ORR is its immunity to methanol crossover and CO poisoning. Chronoamperometric responses to the methanol or CO that were introduced into the  $O_2$ -saturated electrolyte were carried out for the PPY/rGO,  $CN_x$ /G-600 and  $CN_x$ /G-800 catalysts, as shown in Figure 6. After the addition of 0.1 mL methanol at 400 s, the oxygen reduction current for all the catalysts did not show an obvious change. When 50 vol% CO was injected at 400 s, a slight decrease in the current was observed, which may be caused by the decreased partial pressure of  $O_2$  (Henry's law). These results indicate that the  $CN_x$ /graphene composite has high stability and durability towards methanol crossover and CO poisoning.

### 3 Conclusion

In summary, rGO was prepared by hydrothermal treatment and used instead of GO as a substrate to immobilize PPy. In contrast to the agglomeration of the dried PPy/GO compo-

site, the dried PPy/rGO composite is a loose powder and can easily be dispersed in aqueous media. After simple calcination at  $600$ – $800^\circ\text{C}$ , sheet-like  $CN_x$ /graphene with a nitrogen content of 9.8 at%–12.5 at% was obtained. The as-prepared  $CN_x$ /graphene composite has an inherent catalytic activity for the ORR and high stability and immunity towards methanol crossover and CO poisoning. The  $CN_x$ /graphene composite prepared in this way is a new metal-free electrocatalyst for the ORR.

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- 1 Wee J H. Applications of proton exchange membrane fuel cell systems. *Renew Sust Energy Rev*, 2007, 11: 1720–1738
- 2 Kamarudina S K, Achmada F, Daud W R W. Overview on the application of direct methanol fuel cell (DMFC) for portable electronic devices. *Int J Hydrogen Energy*, 2009, 34: 6902–6916
- 3 Gasteiger H A, Markovic N M. Just a dream—or future reality?

- Science, 2009, 324: 48–49
- 4 Rajesh B, Piotr Z. A class of non-precious metal composite catalysts for fuel cells. *Nature*, 2006, 443: 63–66
  - 5 Bjorn W J, Orawan W J, Maria F, et al. High rates of oxygen reduction over a vapor phase-polymerized PEDOT electrode. *Science*, 2008, 321: 671–674
  - 6 Ziegelbauer J M, Gullá A F, O’Laoire C, et al. Chalcogenide electrocatalysts for oxygen-depolarized aqueous hydrochloric acid electrolysis. *Electrochim Acta*, 2007, 52: 6282–6294
  - 7 Bezerra C W B, Zhang L, Lee K C, et al. A review of Fe–N/C and Co–N/C catalysts for the oxygen reduction reaction. *Electrochim Acta*, 2008, 53: 4937–4951
  - 8 Chen J, Zhang W M, Officer D, et al. A readily-prepared, convergent, oxygen reduction electrocatalyst. *Chem Commun*, 2007, 3353–3355
  - 9 Lee K C, Zhang L, Lui H S, et al. Oxygen reduction reaction (ORR) catalyzed by carbon-supported cobalt polypyrrole (Co-PPy/C) electrocatalysts. *Electrochim Acta*, 2009, 54: 4704–4711
  - 10 Lyth S M, Nabae Y, Moriya S, et al. Carbon nitride as a non precious catalyst for electrochemical oxygen reduction. *J Phys Chem C*, 2009, 113: 20148–20151
  - 11 Biddinger E J, Deak D V, Ozkan U S. Nitrogen-containing carbon nanostructures as oxygen-reduction catalysts. *Top Catal*, 2009, 52: 1566–1574
  - 12 Noto V D, Negro E. Development of nano-electrocatalysts based on carbon nitride supports for the ORR processes in PEM fuel cells. *Electrochim Acta*, 2010, 55: 7564–7574
  - 13 Liu G, Li X G, Ganesan P, et al. Development of non-precious metal oxygen-reduction catalysts for PEM fuel cells based on N-doped ordered porous carbon. *Appl Catal B-Environ*, 2009, 93: 156–165
  - 14 Gong K P, Du F, Xia Z H, et al. Nitrogen-doped carbon nanotube arrays with high electrocatalytic activity for oxygen reduction. *Science*, 2009, 323: 760–764
  - 15 Tang Y F, Allen B L, Kauffman D R, et al. Electrocatalytic activity of nitrogen-doped carbon nanotube cups. *J Am Chem Soc*, 2009, 131: 13200–13201
  - 16 Zhu C, Higgins D, Tao H S, et al. Highly active nitrogen-doped carbon nanotubes for oxygen reduction reaction in fuel cell applications. *J Phys Chem C*, 2009, 113: 21008–21013
  - 17 Chen Z, Higgins D, Chen Z W, et al. Nitrogen doped carbon nanotubes and their impact on the oxygen reduction reaction in fuel cells. *Carbon*, 2010, 48: 3057–3065
  - 18 Xiong W, Du F, Liu Y, et al. 3-D carbon nanotube structures used as high performance catalyst for oxygen reduction reaction. *J Am Chem Soc*, 2010, 132: 15839–15841
  - 19 Yu D S, Zhang Q, Dai L M. Highly efficient metal-free growth of nitrogen-doped single-walled carbon nanotubes on plasma-etched substrates for oxygen reduction. *J Am Chem Soc*, 2010, 132: 15127–15129
  - 20 Lefèvre M, Proietti E, Jaouen F, et al. Iron-based catalysts with improved oxygen reduction activity in polymer electrolyte fuel cells. *Science*, 2009, 324: 71–74
  - 21 Qin H Y, Liu Z X, Ye L Q, et al. The use of polypyrrole modified carbon-supported cobalt hydroxide as cathode and anode catalysts for the direct borohydride fuel cell. *J Power Sources*, 2009, 192: 385–390
  - 22 Qin H Y, Lao S J, Liu Z X, et al. Effects of heat treatment on the structure, morphology and electrocatalytic activity of cobalt hydroxide electrocatalyst. *Int J Hydrogen Energy*, 2010, 35: 1872–1878
  - 23 Yoo E J, Okata T, Akita T, et al. Enhanced electrocatalytic activity of Pt subnanoclusters on graphene nanosheet surface. *Nano Lett*, 2009, 9: 2255–2259
  - 24 Seger B, Kamat P V. Electrocatalytically active graphene-platinum nanocomposites. Role of 2-D carbon support in PEM fuel cells. *J Phys Chem C*, 2009, 113: 7990–7995
  - 25 Qu L T, Liu Y, Baek J B, et al. Nitrogen-doped graphene as efficient metal-free electrocatalyst for oxygen reduction in fuel cells. *ACS Nano*, 2010, 4: 1321–1326
  - 26 Shao Y Y, Zhang S, Engelhard M H, et al. Nitrogen-doped graphene and its electrochemical applications. *J Mater Chem*, 2010, 20: 7491–7496
  - 27 Jafri R I, Rajalakshmi N, Ramaprabhu S. Nitrogen doped graphene nanoplatelets as catalyst support for oxygen reduction reaction in proton exchange membrane fuel cell. *J Mater Chem*, 2010, 20: 7114–7117
  - 28 Lee K R, Lee K U, Lee J W, et al. Electrochemical oxygen reduction on nitrogen doped graphene sheets in acid media. *Electrochem Commun*, 2010, 12: 1052–1055
  - 29 Sun Y Q, Li C, Xu Y X, et al. Chemically converted graphene as substrate for immobilizing and enhancing the activity of a polymericcatalyst. *Chem Commun*, 2010, 46: 4740–4742
  - 30 Wei D C, Liu Y Q, Wang Y, et al. Synthesis of N-doped graphene by chemical vapor deposition and its electrical properties. *Nano Lett*, 2009, 9: 1752–1758
  - 31 Li X L, Wang H L, Robinson J T, et al. Simultaneous nitrogen doping and reduction of graphene oxide. *J Am Chem Soc*, 2009, 131: 15939–15944
  - 32 Zhang K, Zhang L L, Zhao X S, et al. Graphene/polyaniline nanofiber composites as supercapacitor. *Chem Mater*, 2010, 22: 1392–1401
  - 33 Cote L J, Kim F, Huang J X. Langmuir Blodgett assembly of graphite oxide single layers. *J Am Chem Soc*, 2009, 131: 1043–1049
  - 34 Jiang X, Ma Y W, Li J J, et al. Self-assembly of reduced graphene oxide into three-dimensional architecture by divalent ion linkage. *J Phys Chem C*, 2010, 114: 22462–22465
  - 35 Guo H L, Wang X F, Qian Q Y, et al. A green approach to the synthesis of graphene nanosheets. *ACS Nano*, 2009, 9: 2653–2659
  - 36 Ma Y W, Jiang S J, Jian G Q, et al. CN<sub>x</sub> nanofibers converted from polypyrrole nanowires as platinum support for methanol oxidation. *Energy Environ Sci*, 2009, 2: 224–229
  - 37 Zhang X T, Zhang J, Song W H, et al. Controllable synthesis of conducting polypyrrole nanostructures. *J Phys Chem B*, 2006, 110: 1158–1165
  - 38 Kudin K N, Ozbas B, Schniepp H C, et al. Raman spectra of graphite oxide and functionalized graphene sheets. *Nano Lett*, 2008, 8: 36–41
  - 39 Choi H C, Song H J, Jung H, et al. Release of N<sub>2</sub> from the carbon nanotubes *via* high-temperature annealing. *J Phys Chem B*, 2005, 109: 1683–1688

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## Supporting Information

**SI 1** Thermogravimetry measurement of PPy/rGO powder

**SI 2** Electrochemical impedance spectroscopy of CN<sub>x</sub>/graphene composite

The supporting information is available online at [csb.scichina.com](http://csb.scichina.com) and [www.springerlink.com](http://www.springerlink.com). The supporting materials are published as submitted, without typesetting or editing. The responsibility for scientific accuracy and content remains entirely with the authors.