

Removal of CO₂ from Flue Gas at High Temperature Using Novel Porous Solids

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Since the CO₂ separation is the first and most energy intensive step of carbon capture and storage (CCS) technology, many research have targeted at improving the current technologies or developing new approaches of CO₂ separation and capture. In this study, lithium orthosilicate-based pellets were developed and characterized as potential regenerable high-temperature CO₂ sorbents. A mechanical method was used for pelletization of the powdered materials, namely K₂CO₃-doped lithium silicate (Li₄SiO₄). For increasing the performance of the pellets over multiple cycles an activation strategy was applied: the powdered sorbents were pelletized with a binder for enhancing their porosity by applying a thermal activation before adsorption process. Different amounts of binders (layered graphite and carbon nanotubes) were added to powered doped-sorbent in order to identify the optimal amount to ensure an adequate porosity into the pellets. The CO₂ sorption properties of the obtained pellets were investigated by using a thermal gravimetric analyzer (TGA) in a controlled gas flow environment at low CO₂ partial pressure (0.04 atm). Compared to the pure K₂CO₃-doped lithium silicate pellets, the sorbents prepared using layered graphite showed greater CO₂ capture capabilities, which were ascribed to the higher porosity developed as a result of the activation. At 580 °C and a CO₂ partial pressure of 0.04 atm, the uptake of CO₂ in pellets prepared with 20 % of graphite reached about 200 mg CO₂/g sorbent within 120 minutes corresponding to a Li₄SiO₄-conversion of 72.1 %. During multiple sorption/desorption cycles, a decay of the sorption capacity of the pellets was observed due to a partial sintering of the materials.

1. Introduction

The emission of carbon dioxide (CO₂), one of the major greenhouse gases, has raised great concerns about the relationship between anthropogenic CO₂ and global warming: CO₂ emissions may have contributed to urban smog, acid rain, and health problems. As a consequence, carbon capture and storage (CCS) technology is attracting increasing interest to reduce the ever-increasing amount of CO₂ released into the atmosphere. CCS involves capturing the carbon dioxide emitted by power plant and other industrial operations and permanently storing it deep underground. On average, implementing a CCS process will increase the capital cost of a power station by 25 to 30 %, and reduce its operating efficiency from about 40 to 30 %. CO₂ capture process accounts for 70-80 % of the overall CCS (Lee and Park, 2015).

CO₂ capture technologies are categorized as post-combustion, pre-combustion, and oxy-fuel combustion. Among these, post-combustion CO₂ capture processes are regarded as being important green and economic technologies (Ahmad et al., 2015). Post-combustion capture uses wet/dry adsorbents, which are used for gas separation, and separates and collects CO₂ by adsorption/desorption. Currently, the commercially available technique for CO₂ capture is based on amine-scrubbing processes (Chavez et al., 2015). However, this technology is energy-intensive, due to the low operating temperature (40-150 °C) and the energy requirement for regeneration of the solutions, and involves environmental penalties, due to the degradation of the expensive amine-based solvents.

Recently, selective adsorption technology (solid sorbents) has been identified as one of the ways for post-combustion CO₂ capture at high temperature, without cooling the flue gas to ambient or even lower temperature (Puccini et al., 2013a). For the separation of carbon dioxide from hot gas stream, the adsorbent

must have high selectivity for CO₂ over N₂ at elevated temperatures and high adsorption capacity for carbon dioxide at the operating temperature.

Among the various solid sorbents that have been studied, lithium orthosilicate (Li₄SiO₄) is considered one of the most promising candidates for the post-combustion CO₂ capture at high temperature (500-700 °C), according to the following reaction r1:



Main advantages of this materials are their high CO₂ capture capacity (theoretically up to 360 mg/g sorbent), lower regeneration temperatures (< 750 °C) compared to other high-temperature sorbents, such as CaO, and their excellent stability that allow operation over a significant number of cycles without losing their sorption capacity.

For large-scale application of such sorbents, the looping cycles system might represent the optimal industrial technology: a typical reactor configuration consist of two interconnected fluidized beds, which allows the circulation of the solid between the reactors and ensures an appropriate contact between the solid and the gas phase, operating alternatively adsorption/regeneration cycles (looping process). Therefore, it is also required that these sorbents are used in the form of granules or pellets. For increasing the performance of the pellets over multiple cycles an activation strategy could be applied: a pelletization procedure of powdered sorbents with a binder (e.g. cellulose fibres) allows enhancing the porosity of pellets by applying a thermal activation before adsorption process (Puccini et al., 2013b).

In this study, pellets of Li₄SiO₄-based sorbents were developed for CO₂ capture at high temperature and low concentration, in view of industrial applications. A mechanical method was used for pelletization of the powdered materials, namely lithium silicate (Li₄SiO₄) doped with binders. The CO₂ sorption properties of the obtained pellets were investigated by using a thermal gravimetric analyzer (TGA) in a controlled gas flow environment at low CO₂ partial pressure (0.04 atm). Scanning electron microscopy (SEM) was used to characterise the prepared sorbents and to understand the relationship between the sorbent structure and CO₂ sorption properties. Besides, the cyclic stability of the pellets was investigated by carrying out multiple cycles of CO₂ sorption/desorption.

2. Experimental

2.1 Sorbent preparation and characterization

Li₄SiO₄ powders were prepared by the solid state method. Starting materials were reagent-grade Li₂CO₃ and SiO₂ (from Sigma-Aldrich) in a 2:1 molar ratio. The materials were weighed, mixed, and ground in an agate mortar with a suitable amount of water. Then, the mixtures were dried and calcined in air at 900 °C for 4 h. Both temperature increase and decrease ramping rates were set to 60 °C/h. The resulting powders were ground by agate mortar to break down any agglomeration and then sieved to produce Li₄SiO₄ particles with size below 10 µm. Promoted Li₄SiO₄-based sorbent, named LiK, was prepared mixing 30 wt% of potassium carbonate (K₂CO₃) with the previously synthesized pure Li₄SiO₄ powder.

Pellets were prepared using the powdered sorbent (LiK) mixed with weighed quantities of binder. Cylindrical pellets, with diameter of 6 mm and length of 2.5 mm, were produced by using a mechanical press, operating at 5 bar pressure applied for 3 min. Different pellet samples were prepared with different binder types and quantities (between 5 and 30 wt%, by weight of LiK). Binder-free pellet was also prepared for comparison purpose (LiK-p).

The sample morphology was observed with a JEOL 5600 LV scanning electron microscope (SEM).

2.2 CO₂ capture tests

Carbon dioxide adsorption on pellets was studied using a thermogravimetric analyzer (TGA Q500 TA Instrument). The pellet was placed in a sample pan and, prior to CO₂ sorption testing, the sample was heated (heating rate of 10 °C/min) in air to required temperature (750 °C) to thermally degrade the binder for developing porosity in the pellets. When the sample weight became stable, the operating temperature (580 °C) was reached, and the N₂ flow was switched to a mixture CO₂/N₂, containing 4 vol% of CO₂. The total feed gas flow rate was maintained at 100 mL/min at atmospheric pressure. The weight increase during the CO₂ sorption process, lasted 120 min, was recorded as a function of time.

Besides, to assess the stability of the optimized sorbent ten adsorption/desorption cycles were carried out. The adsorption of CO₂ was carried for 60 min at 580 °C in 4 vol% CO₂ flow. Afterwards, the feed gas stream was switched to a pure N₂ flow (100 mL/min) and the temperature was raised to 700 °C with heating rate of 20 °C/min. The sample was kept to the regeneration temperature until it was completely regenerated. The regeneration process lasted about 15 min. After that, a new cycle was carried out.

Since the promoters showed no affinity for the CO₂ capture but contributed to the sorbent weight, the conversion of Li₄SiO₄, $X_{\text{Li}_4\text{SiO}_4}$, was used to assess and compare the effect of the different promoters used on the CO₂ sorption reaction. The percentage conversion $X_{\text{Li}_4\text{SiO}_4}$ was calculated by using the Eq(1):

$$X_{\text{Li}_4\text{SiO}_4} (\%) = \frac{\Delta W}{F_{\text{Li}_4\text{SiO}_4} r_s} \quad (1)$$

where ΔW is the percentage weight change of the sorbent during the adsorption, $F_{\text{Li}_4\text{SiO}_4}$ is the weight fraction of Li₄SiO₄ in the sorbent and r_s is the stoichiometric uptake of CO₂ by Li₄SiO₄, which is about 0.36 g of CO₂/g of Li₄SiO₄.

3. Results and discussion

Li₄SiO₄-based pellets were prepared by sintering the powdered sorbent with a binder. A thermal pre-treatment before adsorption process was applied for developing a porous structure into the pellets: during degradation the binder releases gases that alter the porosity of sorbent, and subsequently could enhance CO₂ capture. In a previous work (Puccini et al., 2013b), cellulose fibres was used as binder. In this study layered graphite (LG) and carbon nanotubes (CN) were used for developing a mesoporous structure. The SEM micrograph of the selected binders are reported in Figure 1. The layered graphite showed particles with size around 10 μm . The carbon nanotubes presented the typical structure constituted of fibres having with diameter of 7-15 nm and length of 0.5-10 μm .

Thermogravimetric analysis of binders under air was conducted to assess their thermal degradation behaviour, and then to set the pre-treatment procedure before sorption test. The thermal decomposition profiles of the binders (the mass loss as function of time) are shown in Figure 2. Layered graphite and carbon nanotubes exhibit similar curves: the LG sample showed thermal stability up to 630 $^{\circ}\text{C}$ and a single-step weight loss close to 870 $^{\circ}\text{C}$, and the CN sample is thermally stable up to 730 $^{\circ}\text{C}$ and exhibits a single-step weight loss around 900 $^{\circ}\text{C}$. Both samples reached the complete degradation over 900 $^{\circ}\text{C}$, keeping such temperature for 15 minutes (the selected materials are free ash). Therefore, the pre-treatment procedure was settled with a temperature of 750 $^{\circ}\text{C}$ (heating rate of 10 $^{\circ}\text{C}/\text{min}$) lasting 30 minutes.

The SEM micrographs of pellet without binder, and the samples after the thermal activation are reported in Figure 3. The SEM images display the morphology of the section of the samples: the picture (a), pellet LiK-p, presents a compact structure without cavities, instead in the picture (a) and (b) holes and cavity are visible, due to the degradation of binders during the thermal pre-treatment.

Different ratios of binders were used in order to identify the optimal amount to ensure good hardness and a porous structure to the pellets. The tested percentage of LG was varied between 20 and 40 wt%. An amount of binder exceeding 40 wt% has not been considered appropriate since reduces the amount of active sorbent. Since problems in the production of pellets during mechanical compression of the powder containing the carbon nanotubes were occurred, the tested percentage of CN was varied between 5 and 15 wt%.

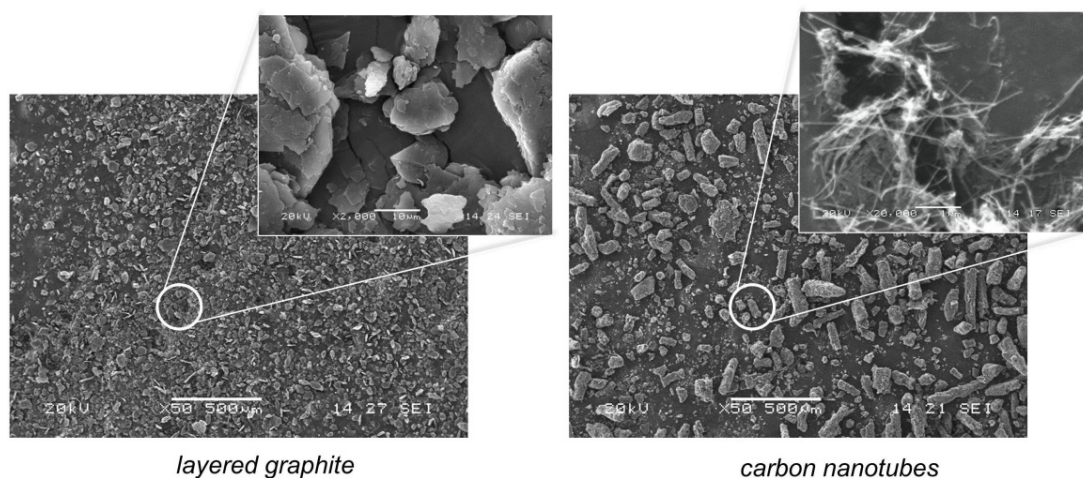


Figure 1: SEM micrograph of binders.

The Figures 4 shows the CO₂ uptake of the pellets as a result of tests in 4 vol% CO₂ stream at temperature of 580 °C. The profile of the cycle displays a typical solid–gas reaction profile. During carbonation, the low resistance diffusion of CO₂ through the pores enabled a fast reaction stage, followed by a slow reaction stage related to the closure of small pores and full masking of surface area by the product of carbonation, i.e. CaCO₃. The Table 1 reports the percentage conversion of the sorbents.

As shown, the LiK sample and the pellets containing carbon nanotubes exhibit similar CO₂ uptake profiles: the adsorbed amount of CO₂ after 2 hours was nearly 0.09 g CO₂/g sorbent corresponding to a 30 % conversion of Li₄SiO₄. The slow reaction stage where both profiles exhibited similar slopes suggests nearly identical diffusion rates of carbon dioxide in both sorbents, therefore a similar morphology of the pellets. The use of carbon nanotubes developed a structure with pores smaller than 1 μm that the CaCO₃ produced during carbonation could be blocked, increasing the resistance to CO₂ diffusion.

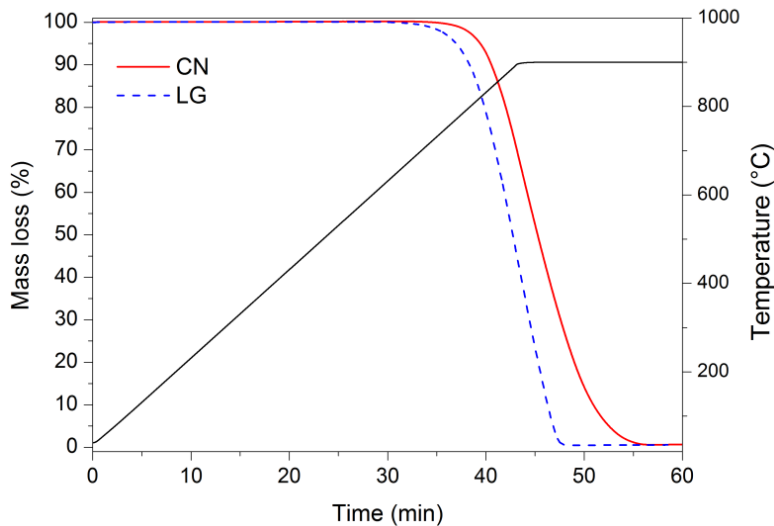


Figure 2: Thermal decomposition profiles of binders (in air and heating rate of 20 °C/min).

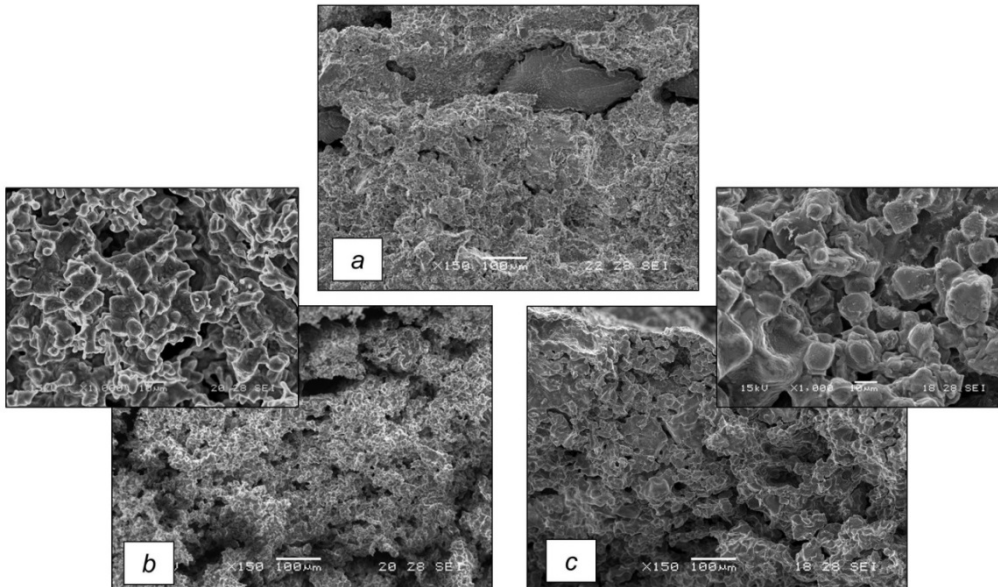


Figure 3: SEM micrograph of cross-section of LiK pellets: (a) without binder, after activation (b) with 20 wt% of layered graphite, (c) with 15 wt% of carbon nanotubes.

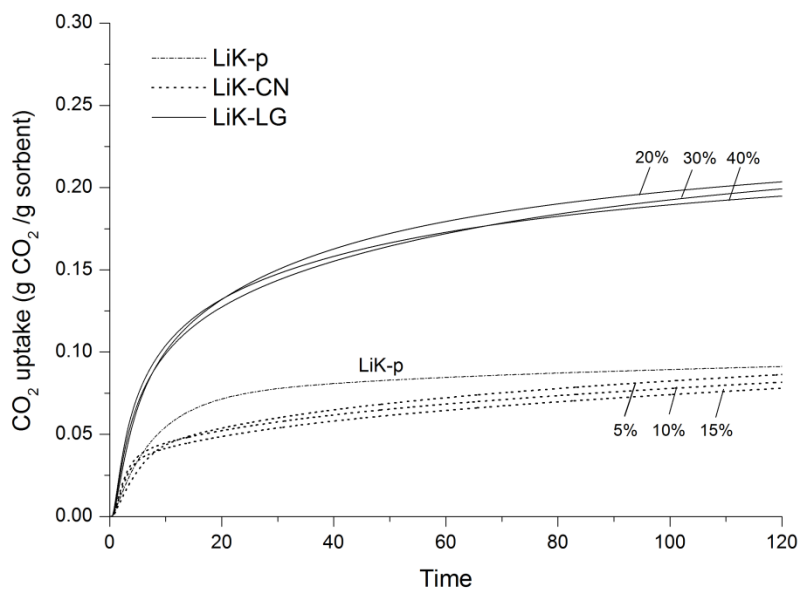
Table 1: Percentage conversion of Li_4SiO_4 .

| Pellets LiK | | $X_{\text{Li}_4\text{SiO}_4}$ (%) |
|-------------|-----|-----------------------------------|
| Binder | wt% | |
| LG | 20 | 72.1 |
| LG | 30 | 69.1 |
| LG | 40 | 57.7 |
| CN | 5 | 30.6 |
| CN | 10 | 27.1 |
| CN | 15 | 28.9 |

Instead, the use of layered graphite enhanced the CO_2 uptake: the adsorbed amount of CO_2 after 2 hours was nearly $0.2 \text{ g CO}_2/\text{g}$ sorbent corresponding to a 72 % conversion of Li_4SiO_4 for the sample produced with 20 % of LG. No significant variation on CO_2 adsorption capacity was observed varying the binder content. This result can be attributed to the different morphology of the pellets: the porosity created as a result of the graphite degradation increased the surface area of the pellets by developing larger pores than CN, and, consequently, their sorption capacity increased.

The cyclic stability of the pellets was investigated by carrying out 10 cycles of CO_2 sorption/desorption (Figure 5). The pellet produced without cellulose fibres (LiK) showed a remarkable decay in CO_2 absorption capacity after the first cycle: the conversion dropped off from 32 % to 13 %, and remained essentially constant during the subsequent cycles. The other pellet, produced using LiK containing the 20 % of layered graphite, showed a different behaviour: the sorption capacity decreased during first four cycles and remained nearly constant for cycles 5-10 (around a 45 % conversion of Li_4SiO_4).

sorption strongly varies and remains almost constant. The decrease of CO_2 uptake with increasing number of cycles demonstrates that the structure of Li_4SiO_4 -base pellets experiences sintering typical of such sorbents, e.g. excessive loss of narrow pores useful for carbonation (Fennell et al., 2007).

Figure 4: CO_2 uptake in 4 vol% CO_2 stream at $580 \text{ }^\circ\text{C}$.

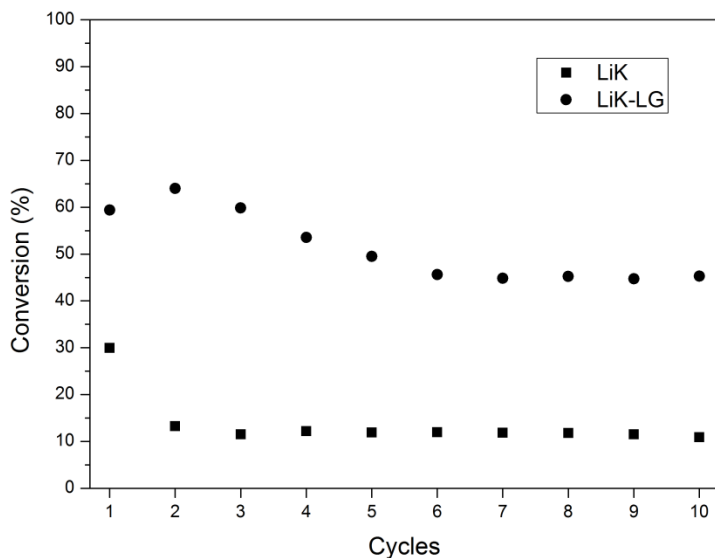


Figure 5: Multiple cycles of CO₂ sorption desorption (in 4 vol% CO₂ stream at 580 °C) on pellets LiK and LiK with 20 wt% of layered graphite

4. Conclusions

In this study novel Li₄SiO₄ based-sorbents in form of pellets for CO₂ capture at high temperature (500 - 600 °C) were developed. A mechanical method was used for pelletization of the powdered materials, namely lithium silicate (Li₄SiO₄) doped with carbonates. Binders were used to make porous the structure as evidenced by the morphology of the produced pellets. The CO₂ sorption properties of the pellets were investigated by using a thermal gravimetric analyzer (TGA). The curves uptake revealed that the use of layered graphite enhanced the CO₂ sorption capacity of the Li₄SiO₄-based pellets. At 580 °C and under a 4 vol% CO₂ flow the CO₂ sorption capacity for the pellets (diameter 6 mm and length 3.5 mm) produced using 20 wt% of graphite was about 200 mg CO₂/g sorbent after 120 minutes corresponding to a Li₄SiO₄-conversion of 72.1 %. These results can be attributed to the morphology of the pellets: the porosity created by the graphite degradation in air (before the first sorption run) increased the surface area of the pellets and, consequently, their sorption capacity.

During 10 CO₂ sorption/desorption cycles, the sorption capacity of the pellets remained nearly constant only for the first 5 cycles and, then, decreased markedly because of a partial sintering of the material, as shown by SEM morphological analysis. Future work will focus on the use of an appropriate inert binder for the pelletization process in order to improve the cyclic stability and the mechanical resistance of the produced pelletized Li₄SiO₄-based sorbents.

Reference

- Ahmad M.Z., Hashim H., Yunus N.A., Muis Z.A., 2015, Optimal solvent design for CO₂ capture process, *Chemical Engineering Transactions*, 45, 1135-1140
- Chavez R.-H., Guadarrama J.J., 2015, Numerical evaluation of co2 capture on post-combustion processes, *Chemical Engineering Transactions*, 45, 271-276
- Fennell S.P., Pacciani R., Dennis J. S., Davidson J. F., Hayhurst A. N., 2007, The effects of repeated cycles of calcination and carbonation on a variety of different limestones, as measured in a hot fluidized bed of sand. *Energy & fuels*, 21(4), 2072-2081
- Lee S., Park S., 2015, A review on solid adsorbents for carbon dioxide capture, *Journal of Industrial and Engineering Chemistry*, 23, 1-11
- Puccini M., Seggiani M., Vitolo S., 2013a, CO₂ capture at high temperature and low concentration on Li₄SiO₄ based sorbents, *Chemical Engineering Transactions*, 32, 1279-284
- Puccini M., Seggiani M., Vitolo S., 2013b, Lithium silicate pellets for CO₂ capture at high temperature, *Chemical Engineering Transactions*, 35, 373-378