“Initial Potential” Effect on the Dissociative Adsorption of Methanol on A Roughened Platinum Electrode in Acidic Solution

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Abstract: In situ Raman spectroscopic and voltammetric studies indicate that dissociative adsorption of methanol on the rough platinum electrode occurs in the hydrogen ad/desorption potential range, and the dissociative extent depends on the initial potential of the electrode before contacting methanol, in addition to the contacting time. As the dissociative product, carbon monoxide competes the site of strongly bound hydrogen preferentially, and shifts the ad/desorption potentials of weakly bound hydrogen towards more positive ones gradually with the increase of CO coverage. Whereas, formaldehyde dissociates more easily by far and completely suppresses H-adsorption. The confocal Raman spectroscopy developed on transition metals shows some intriguing advantages in investigating electrocatalytic oxidation of small organic molecules.

Keywords: Methanol, formaldehyde, dissociative adsorption, in situ Ramam spectroscopy.

The research of electrocatalytic oxidation of small organic molecules such as methanol, formaldehyde and formic acid has been fascinating electrochemists for more than thirty years, owing to its significance in energy conversion and reaction mechanisms. The surface adsorbed species involved in these electrocatalytic reactions, especially reactive and poisoning intermediates, have been investigated by electrochemical, spectroscopic and other physicochemical methods1. However, Raman spectroscopy has hardly been used as a technique so far to study the electrocatalytic oxidation processes of these reactants on Pt electrodes mainly due to its low sensitivity for the surface adsorbed species from that electrode2. With an appropriate surface roughening procedure for bare Pt electrodes3 to gain the surface-enhanced Raman spectroscopic (SERS) effect and by using a confocal Raman microscope (LabRam I, Dilor, France)4, we are able to obtain good quality surface Raman signals of adsorbed CO as the dissociative product of small organic molecules on rough Pt electrodes with a thin-layer of solution as thick as 1 mm, which allows a large Faradaic current without hampering the electrochemical reactions.

In this work, the dissociative adsorption of methanol was studied by in situ surface Raman spectroscopy as well as electrochemical cyclic voltammetry. Some interesting results were obtained while different potentials were imposed on the electrode before the reactant was added into the blank solution, 1 mol/L H2SO4, which is what we mean by an “initial potential”. A comparative study was performed with formaldehyde.
The roughened electrode\textsuperscript{3} underwent a potential cycling between –0.2~1.2 V vs. SCE in the blank solution of 1 mol/L H\textsubscript{2}SO\textsubscript{4} till a repeatable cyclic voltammogram appeared (curve a’s in Figure 1). Then methanol was admitted in while the cell was either on (controlling different initial potentials, –0.2, –0.1 V) or open, and Raman measurement followed.

Vibrational bands of 490 and 2050 cm\textsuperscript{-1} in Raman spectra of Figure 2 are from the Pt–C and C≡O of the linearly bound CO respectively, and the ability to obtain both the low frequency and the high frequency vibrations of Pt–C and C≡O bands allows the assignment of surface species unambiguously. The spectra show that the dissociative adsorption of methanol can occur at –0.1 V (Figure 2b), as well as at open circuit potential (Figure 2c), but doesn’t happen at –0.2 V (Figure 2a). The open circuit potential (OCP) of the platinum electrode in the blank solution was ca. 0.66 V, and it drifted to ca. 0.2 V in the methanol-containing solution, a potential value at which methanol dissociates fairly readily.

Different dissociative behavior was also observed with the variation of an upper potential limit of the cyclic voltammetry under the same initial potential control of –0.2 V. For example, a new dissociative adsorption peak appeared around 0.2 V (curve b of Figure 1A) during the first forward potential scan between –0.2~0.6 V, and then the
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Figure 2  Surface Raman spectra

On the same Pt electrode of Figure 1 in 1 mol/L H₂SO₄ with (a–c) 0.5 mol/L CH₃OH, (d) 0.5 mol/L HCHO, while the cell was either (c) open or on with an initial potential control at (a) −0.2 V, (b) −0.1 V, (d) −0.2 V.

Neither hydrogen ad/desorption peaks nor the new peak could be observed while the cell was open before adding the methanol. So a much higher oxidative desorption peak (i.e., a larger area) from the adsorbed CO at ca. 0.5 V (Figure 1D) was observed, when the electrode was initially at OCP than at −0.1 V, with the same period of contacting time of 5 min in the methanol-containing solution before its transferring into
the blank solution. Their area ratio for the oxidative desorption of the adsorbed CO in Figure 1D is basically consistent with that for the vibrational band intensity of C≡O or Pt-C.

Unlike methanol, which does not dissociate on the electrode surface at an initial potential of −0.2 V (Figure 2a), formaldehyde dissociates much more easily on the platinum electrode than methanol does, and the hydrogen ad/desorption peaks are completely suppressed by the adsorbed CO even for the first potential cycle (not shown), which is consistent with the Raman spectra in Figure 2d, where the band intensity of Pt–C and C≡O vibrations at −0.2 V is comparable with that for methanol at OCP (ca. 0.2 V). The difference in dissociative behavior may be due to different dissociative mechanisms as suggested by Parsons.5

Through this preliminary study, we have demonstrated that the availability of confocal Raman microscopy offers intriguing opportunities for investigating surface reaction on transition metals in electrochemical environments with at least following advantages: direct vibrational spectra can be obtained on rough electrodes with dark color, under high Faradaic currents, with high sensitivity at low frequency vibrations and low interference from bulk solution. Further investigation is now underway in our groups.

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References


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