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<th>Phase Stability of Bi-2(V₁₋ₓMₓ)O₅.₅₊δ (Mₓ = Li and Ag, x = 0.05 and 0.1)</th>
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<td>Author(s)</td>
<td>Taninouchi, Yu-ki; Uda, Tetsuya; Ichitsubo, Tetsu; Awakura, Yasuhiro; Matsubara, Eiichiro</td>
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Phase Stability of Bi$_2$(V$_{1-x}$ME$_x$)O$_{5.5+x}$ (ME = Li and Ag, x = 0.05 and 0.1)  
Yu-ki Taninouchi*, Tetsuya Uda, Tetsu Ichitsubo, Yasuhiro Awakura and Eiichiro Matsubara  
Department of Materials Science and Engineering, Kyoto University, Kyoto 606-8501, Japan  

The phase transition of high oxide-ion conductor Bi$_2$(V$_{1-x}$ME$_x$)O$_{5.5+x}$ (ME = Li and Ag, x = 0.05 and 0.1) and its long-term phase stability against thermal decomposition have been studied. The transition behavior was determined by differential scanning calorimetry (DSC) as well as high-temperature X-ray diffraction (HT-XRD) analysis. The observed thermal and structural changes during temperature scans reveal the temperature dependence of electrical conductivity. Li- and Ag-doping do not sufficiently suppress thermal decomposition at intermediate temperatures between 400 and 600°C. In particular, partial decomposition was detected in Bi$_2$(V$_{0.95}$Li$_{0.05}$)O$_{5.4}$ during a temperature scan by HT-XRD analysis, which explains the distinct change in electrical conductivity. We have generated pseudo-binary phase diagrams for Bi$_2$(V$_{1-x}$ME$_x$)O$_{5.5+x}$ based on the X-ray diffraction analysis of powders annealed for 200 h and differential thermal analysis. The thermodynamically stable region of Bi$_2$(V$_{1-x}$ME$_x$)O$_{5.5+x}$ is not sufficiently expanded by Li- and Ag-doping. At around 500°C, Bi$_2$(V$_{1-x}$ME$_x$)O$_{5.5+x}$ is metastable, although it shows the highest oxide-ion conductivity among solid oxides.  


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Keywords: BIMEVOX, Bi$_2$VO$_{5.5}$, phase transition, phase stability, phase decomposition, lithium, silver  

1. Introduction  

Bi$_2$(V$_{1-x}$ME$_x$)O$_{5.5+x}$ (ME: dopant metal, x: dopant concentration, δ: deviation of oxygen content from original stoichiometry), i.e., BIMEVOX, is a strong candidate for oxide-ion conduction at intermediate temperatures between 400 and 600°C. Non-doped Bi$_2$VO$_{5.5}$ has a single-layer Aurivillius structure and shows three main polymorphs: α (below 450°C), β (from 450 to 570°C), and γ (above 570°C). Their lattice structures are characterized by an orthorhombic subcell of $a_\text{m} = 0.553$, $b_\text{m} = 0.561$, and $c_\text{m} = 1.528$ nm, i.e., $a \approx 3a_\text{m}$, $b \approx b_\text{m}$, $c \approx c_\text{m}$, $\gamma \approx 90.26^\circ$ for monoclinic $a$, $a \approx 2a_\text{m}$, $b \approx b_\text{m}$, $c \approx c_\text{m}$ for orthorhombic $\beta$, and $a \approx a_\text{m}/\sqrt{2}$, $c \approx c_\text{m}$ for tetragonal $\gamma$. Characteristic superlattice diffraction peaks are observed in the diffraction patterns of α- and β-Bi$_2$VO$_{5.5}$, and γ-Bi$_2$VO$_{5.5}$ exhibit oxide-ion conduction. In particular, the conductivity of γ-phase is as high as $3 \times 10^{-11}$ S cm$^{-1}$.  

The polymorphs of Bi$_2$VO$_{5.5}$ and their electrical properties are significantly affected by doping. We have reported that the γ-phases of monovalent-metal-doped Bi$_2$VO$_{5.5}$ [Bi$_2$(V$_{1-x}$ME$_x$)O$_{5.5+x}$; ME = Li and Ag, x = 0.05 and 0.1] are stabilized below 570°C and show quite high oxide-ion conductivities at around 500°C. Their oxide-ion conductivities are $0.5-1 \times 10^{-11}$ S cm$^{-1}$ at 500°C, which are the highest values among solid oxides, and almost the same order of those of (ZrO$_2$)$_{0.02}$(V$_2$O$_5$)$_{0.08}$ at around 850°C, (CeO$_2$)$_{0.85}$(SmO$_{1.5}$)$_{0.15}$ at around 750°C, and of La$_{0.85}$Sr$_{0.15}$Ga$_2$O$_3$ at around 600°C.  

In the present study, the phase transitions of Li- and Ag-doped Bi$_2$VO$_{5.5}$ were investigated by differential scanning calorimetry (DSC) and high-temperature X-ray diffraction (HT-XRD) analysis in order to understand the change in the electrical conductivity. Recently, Watanabe$^{21}$ reported that Bi$_2$VO$_{5.5}$ is metastable below 550°C and it decomposes to BiVO$_4$ and Bi$_5$V$_{12}$O$_{25}$ for 300 h or more. Steil et al.$^{13}$ also confirmed that Bi$_2$VO$_{5.5}$ completely decomposes into BiVO$_4$ and Bi$_5$V$_{12}$O$_{25}$ at 500°C for 7 weeks. The long-term phase stability is the key for the practical use. Therefore, we assessed the phase stabilities of Li- and Ag-doped Bi$_2$VO$_{5.5}$ and established the phase diagrams.  

In this paper, the phases are simply classified as α-, β-, or γ-phase according to the characteristic superlattice diffraction peaks, irrespective of the symmetry of the unit cell. That is, the orthorhombic phases of Bi$_2$(V$_{0.95}$Li$_{0.05}$)O$_{5.4}$ and Bi$_2$(V$_{0.95}$Ag$_{0.05}$)O$_{5.4}$ that show characteristic superlattice diffraction peaks of α-Bi$_2$VO$_{5.5}$ are classified as α-phases. The tetragonal phases that show characteristic superlattice diffraction peaks of β-Bi$_2$VO$_{5.5}$ are classified as β-phases.  

2. Experimental  

The polycrystalline powders listed in Table 1 were synthesized by the solid-state reaction in air using platinum containers.$^8$ Appropriate amounts of Bi$_2$O$_3$ (99.99%; Furuchi Chemical), V$_2$O$_5$ (99.9%; Furuchi Chemical), Li$_2$CO$_3$ (99.9%; Furuchi Chemical), and Ag$_2$O (99%+; Wako) were mixed and then heated at 600°C for 12 h. The ground powders were pelletized at 19.6 MPa, heated again at 800°C (Li$_5$, Ag$_5$, and Ag$_{10}$) or 700°C (Li$_{10}$) for 12 h. and then ground. Before analysis, synthesized powders were kept at 700°C (Li$_5$, Li$_{10}$, and Ag$_5$) or 750°C (Ag$_{10}$) for several hours and then quenched in air.  

DSC was performed using a Diamond differential scanning calorimeter (PerkinElmer) with platinum containers. Powders were identified at room temperature from powders quenched from 700°C (Li$_5$, Li$_{10}$, and Ag$_{10}$) or 750°C (Ag$_{10}$).$^8$  

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<th>Abbreviation</th>
<th>Composition</th>
<th>Phase at room temperature</th>
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<tr>
<td>Li5</td>
<td>Bi$<em>2$(V$</em>{0.95}$Li$<em>{0.05}$)O$</em>{5.4}$</td>
<td>orthorhombic α</td>
</tr>
<tr>
<td>Li10</td>
<td>Bi$<em>2$(V$</em>{0.95}$Li$<em>{0.05}$)O$</em>{3.3}$</td>
<td>tetragonal γ</td>
</tr>
<tr>
<td>Ag5</td>
<td>Bi$<em>2$(V$</em>{0.95}$Ag$<em>{0.05}$)O$</em>{4.4}$</td>
<td>orthorhombic α</td>
</tr>
<tr>
<td>Ag$_{10}$</td>
<td>Bi$<em>2$(V$</em>{0.95}$Ag$<em>{0.05}$)O$</em>{3.3}$</td>
<td>tetragonal γ</td>
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*Graduate Student, Kyoto University, Corresponding author, E-mail: taninouchi@t01.mbox.media.kyoto-u.ac.jp
ders were heated and cooled at 10 K min\(^{-1}\) between 100 and 700°C in artificial air. HT-XRD analysis was carried out in air using X’Pert PRO MPD (PANalytical, Cu-K\(\alpha\) radiation) with an HTK-1200N oven chamber (Anton Paar). The scan rate of the sample temperature was 10 K min\(^{-1}\). The sample temperatures were maintained for 5 min at 50°C intervals, and diffraction patterns were collected within 10 min at each temperature.

The long-term phase stability was assessed by carrying out XRD analysis of the powder samples annealed at various temperatures for 200 h in air. Compositions of the annealed powders were determined by measuring the concentration of metallic ions by energy-dispersive X-ray microanalysis (EDX) using JSM-6500F (JEOL), on the assumption that oxygen content follows the stoichiometry evaluated from the valences of the metallic ions. Differential thermal analysis (DTA) was also carried out using DTG-60H (Shimadzu) until the liquid phase appeared. Powders in alumina containers (DTA) was also carried out using DTG-60H (Shimadzu) until the liquid phase appeared. Powders in alumina containers were heated to 10 K min\(^{-1}\), and diffraction patterns were collected within 10 min at each temperature.

3. Results and Discussion

3.1 Phase transitions and phase decompositions determined by DSC and HT-XRD analysis

Figure 1 shows the DSC profiles of Li- and Ag-doped Bi\(_2\)VO\(_{5.5}\) with the reference profile of Bi\(_2\)VO\(_{5.5}\). As seen in Fig. 1, Bi\(_2\)VO\(_{5.5}\) shows two endothermic peaks on heating and two exothermic peaks on cooling, corresponding to the transitions between \(\alpha\)- and \(\beta\)-phases and \(\beta\)- and \(\gamma\)-phases, respectively. Figure 2 shows the HT-XRD profiles on heating, and Fig. 3 shows a summary of the phases confirmed by HT-XRD analysis. We must note that the diffraction intensities shown in Fig. 2 are plotted in the logarithmic scale to clarify the superlattice diffraction peaks and small peaks of secondary phases.

The transition between \(\alpha\)- and \(\gamma\)-phases passes through the \(\beta\)-phase in Li5 and Ag5. In the DSC profiles of Li5 and Ag5 on heating, clear endothermic peaks appeared from 400 to 450°C, corresponding to the transition from \(\alpha\)- to \(\beta\)-phases (\(\alpha/\beta\) transition). It should be noted that the \(\beta\)-phases of Li5 and Ag5 at 450°C on heating have tetragonal symmetry, as shown in Fig. 2(a) and (c). Very broad and small endothermic peaks were observed for the \(\beta/\gamma\) transition at approximately 480°C in DSC profiles. The broad peak can be explained on the basis of the nature of the second-order \(\beta/\gamma\) transition in Li5 and Ag5. DSC profiles are consistent with continuous conductivity changes at the \(\beta/\gamma\) transition, as shown in Fig. 4. In the DSC profiles of Li5 and Ag5 on cooling, the exothermic peaks at around 420°C correspond to the \(\gamma/\beta\) transition. The exothermic peaks corresponding to the \(\beta/\alpha\) transition are very small and broad, which indicates that the \(\beta/\alpha\) transitions in Li5 and Ag5 proceed gradually.

In Li10 and Ag10, the fundamental diffraction peaks were attributed to the \(\gamma\)-phase. In the DSC profile of Li10, two endothermic peaks are exhibited in the temperature range from 300 to 420°C. In the same temperature range, the diffraction peaks were broadened as shown in Fig. 3(b) and the conductivity increased in Fig. 4. Considering that the broadened peaks split in the orthorhombic \(\alpha\)-phase and became a single peak in the tetragonal \(\beta\)- and \(\gamma\)-phases, as shown in Fig. 2(a), the broadening of the peaks at 300 and 350°C in Fig. 2(b) is attributed to some structural relaxation induced by the rearrangement of oxygen vacancies and/or dopant Li. A pair of endothermic and exothermic peaks is observed at around 680°C in the DSC profile of Li10, as indicated by asterisks. These peaks correspond to a partial melting of Li10, as explained in the following section. In the case of Ag10, no clear peak is observed in the DSC profiles, although Ag10 shows a slow conductivity jump from 300 to 400°C, as shown in Fig. 4.

Results of the HT-XRD analysis also showed the progress of thermal decompositions in Li10, Ag5, and Ag10 during the temperature scan. In Li10, a very small amount of the secondary phase appeared from 550 to 650°C on heating. Similarly, a small amount of the secondary phase appeared in Ag5. The results are summarized in Fig. 3. In Ag10, a relatively large amount of the secondary phase appeared from 550 to 700°C on heating, as shown in Fig. 2(d). The composition of the secondary phase is estimated to be Bi\(_{12}\)VO\(_5\)Ag\(_{28}\). The ratio of the maximum peak of the secondary phase at \(\sim 27.1^\circ\) to that of BIMEVOX at \(\sim 28.2^\circ\), \(I_\beta/I_\gamma\), is shown in Fig. 5 as a function of temperature. The amount of secondary phase increases from 550 to 650°C and then decreases from 650 to 700°C. Thus, the distinct change in the conductivity of Ag10 from 540 to 730°C, which is shown in region (i) of Fig. 4, is explained by the decomposition reaction from 540 to 630°C,

\[\text{Ag10} \rightarrow \text{BIMEVOX phase} + \text{secondary phases},\]

and the recombination reaction from 630 to 730°C,
BIMEVOX phase + secondary phases → Ag10.

On cooling, a small amount of the secondary phase \( (I_{SP}/I_{BVO}, \sim 1.5\%) \) appeared below 550°C in Ag10.

The appearance of secondary phases during the HT-XRD analysis indicates the instability of Li- and Ag-doped Bi\(_2\)VO\(_5\) at intermediate temperatures. Li and Ag have different ionic radii,\(^{14}\) Pauling electronegativities,\(^{15}\) and preferable oxygen coordinations, as summarized in Table 2. However, the transition behavior does not show a large difference between Li- and Ag-doped Bi\(_2\)VO\(_5\) at the same dopant concentration. The dopant concentration changes the oxygen content of BIMEVOX to preserve electroneutrality. Thus, the dominant factors that determine the transition behavior of BIMEVOX are dopant and oxygen contents.

3.2 Long-term phase stabilities and phase diagrams determined by XRD analysis of annealed powders and DTA

Figure 6 shows the XRD patterns of powders after annealing for 200 h at 500 or 700°C. At 500°C, Bi\(_2\)VO\(_5\) is partially decomposed during annealing. The decomposition was also confirmed in doped Bi\(_2\)VO\(_5\) at 500°C. In particular, the BIMEVOX phase almost disappeared in annealed Li5. Ag10 annealed at 500°C contained the BIMEVOX phase, BiVO\(_3\), and two other phases whose compositions are roughly evaluated as Bi\(_{14}\)V\(_3\)AgO\(_{38}\) and Bi\(_3\)V\(_2\)AgO\(_{10}\) by EDX. In Ag10 annealed at 700°C, a secondary phase exists although the BIMEVOX phase remained as the main phase. EDX indicated that the composition of the secondary phase was Bi\(_{14}\)V\(_3\)AgO\(_{38}\), and the composition of the BIMEVOX phase remained largely...
unchanged as Bi$_2$(V$_{0.9}$Ag$_{0.1}$)O$_{5.3}$. Therefore, it can be concluded that annealed Ag10 is a mixture of the BIMEVOX phase and Bi$_2$V$_5$AgO$_{34}$ at 700°C.

Similar measurements were performed at various temperatures. The results are summarized as pseudo-binary phase diagrams in Fig. 7. Li- and Ag-doped Bi$_2$VO$_{5.5}$ are meta-

stable at intermediate temperatures. At 400°C, thermal decompositions were kinetically suppressed in Li10 and Bi$_2$VO$_{5.5}$, and the ordering from γ- to α-phase progressed in Li10. The upper limit of the stable region of the BIMEVOX phase was evaluated by DTA. Partial melting of Bi$_2$VO$_{5.5}$ commenced at 860°C. Figure 8 show the DTA profile of Li10.
on heating as an example. A slight endothermic reaction started at $T_1 = 680\, ^\circ\text{C}$. At the same temperature, an endothermic peak was observed in the DSC profile, as shown in Fig. 1. The DTA profile also shows a large endothermic peak, which appeared at $T_2 = 855\, ^\circ\text{C}$, and disappeared at $T_3 = 887\, ^\circ\text{C}$. Only the BIMEVOX phase was observed in the XRD analysis after annealing for 200 h as well as DTA on heating to the temperature at which the liquid phase forms. The shaded areas indicate the thermodynamically stable region of the BIMEVOX phase. An example of a DTA profile is shown in Fig. 8. $T_1$, $T_2$, and $T_3$ represent the onset of a slight endothermic reaction, the onset of a large endothermic reaction, and the end of an endothermic reaction, respectively. Estimated phases in each region are as follows: (i) BIMEVOX phase + small amount of liquid, (ii) solid + liquid, (iii) liquid.

4. Conclusion

In this study, we investigated the phase transitions and long-term phase stabilities of Bi$_2$(V$_{1-x}$ME$_x$)O$_{5.5+\delta}$ (ME = Li and Ag, $x = 0.05$ and 0.1). Our findings are summarized as follows:

(1) The phase transition behavior is revealed by DSC and HT-XRD analysis. For example, Bi$_2$(V$_{0.95}$Li$_{0.05}$)O$_{5.4}$ and Bi$_2$(V$_{0.95}$Ag$_{0.05}$)O$_{5.4}$ show transitions from $\alpha$- to $\beta$-phase and $\beta$- to $\gamma$-phase, which explains the changes in electrical conductivity. Both types of 5 mol% doped Bi$_2$VO$_{5.5}$ showed the similar phase transitions.

(2) The Li- and Ag-doping do not improve the phase stability against thermal decomposition. In Bi$_2$(V$_{0.9}$Ag$_{0.1}$)O$_{5.3}$, thermal decomposition is easily occurred even during HT-XRD analysis, which explains the distinct change in electrical conductivity. On the basis of the XRD analysis after long-term annealing for 200 h, Bi$_2$(V$_{1-x}$ME$_x$)O$_{5.5+\delta}$ is confirmed to be metastable at around 500°C.

(3) We generated a pseudo-binary phase diagram as a function of dopant concentration. The stable region of the BIMEVOX phase is reduced by Li- and Ag-doping. At around 500°C, Li- and Ag-doped Bi$_2$VO$_{5.5}$ show the highest oxide-ion conductivities among solid oxides. Practical use of these compounds is, however, difficult owing to poor thermal stabilities.

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REFERENCES