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Kyoto University
Crystallization Behavior and Structural Stability of Zr\textsubscript{50}Cu\textsubscript{40}Al\textsubscript{10} Bulk Metallic Glass

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We have investigated the origin of significantly high thermal stability of Zr\textsubscript{50}Cu\textsubscript{40}Al\textsubscript{10} metallic glass and its crystallization behavior as compared to Zr\textsubscript{70}Cu\textsubscript{20}Al\textsubscript{10} and Zr\textsubscript{50}Cu\textsubscript{30} glassy alloys, by differential scanning calorimetry (DSC), X-ray diffraction (XRD) and resonant ultrasound spectroscopy (RUS) techniques. It was found from XRD and DSC analyses that (i) constituent atoms in Zr\textsubscript{50}Cu\textsubscript{40}Al\textsubscript{10} are highly close-packed in a glassy state and (ii) Zr\textsubscript{50}Cu\textsubscript{40}Al\textsubscript{10} glass shows a complicated crystallization process when it is heated at conventional heating rates, but when it is rapidly heated, the ordered B2-type ZrCu is dominantly formed. The high thermal stability and its crystallization behavior are discussed in terms of the long-range-diffusion of atoms and elastic strain energy.

1. Introduction

Bulk metallic glasses, so-called “BMG”s, exhibiting a wide supercooled liquid region have attracted considerable attention not only from engineering aspects but also from glass science perspectives.\textsuperscript{1} They are thermally stable against crystallization, despite the fact that they are mainly composed of relatively-weak metallic bonds and distinct from major glasses having network structures consisting of covalently-bonded clusters or molecules. Inoue’s “three experimental rules”\textsuperscript{2} are very helpful to predict and produce metallic glasses with high glass-forming ability (GFA), but there are still unsolved problems on the thermal stability, and many researchers have tried to elucidate the origin of the excellent thermal stability of BMGs.

In our opinion, the most important thing in the practical viewpoints is to understand how BMGs can be prevented from crystallization upon heating or thermal process, that is, how the thermal stability of BMGs is enhanced. Here we are now considering four effects for enhancement of the thermal stability of BMGs: (i) closely packing of constituent atoms,\textsuperscript{3} (ii) structural dissimilarity between the amorphous phase and crystallized phase(s),\textsuperscript{4} (iii) necessity of the long-range-diffusion of atoms during crystallization, and (iv) a large amount of the elastic strain energy yielded by crystallization in a glassy solid. In the Zr\textsubscript{70}Cu\textsubscript{30} amorphous alloy, for example, the distance of Cu-Cu pairs in amorphous alloy is distinctly different from that in the stable crystalline phase Zr\textsubscript{2}Cu\textsubscript{3},\textsuperscript{5,6} Referring to the structural-dissimilarity aspect, the significant difference in Cu-Cu distance would prevent the crystallization and lead to a stable glass transition. By contrast, crystallization easily occurs in the Zr\textsubscript{70}Ni\textsubscript{30} amorphous alloy because of the structural similarity. In addition, the thermal stability of Zr-Cu-Al alloys is markedly varied by the addition of Al element.\textsuperscript{7,8} As discussed later, this is mainly because the atomic diffusion of Al atoms is required to form thermally stable crystalline phases. Moreover, the strain-energy concept is a new one for obtaining a stable glass-to-liquid transition, which will be deeply discussed in this paper.

In general, the glass transition temperature (\(T_g\)) increases only slightly with increase in a heating rate, whereas the crystallization temperature (\(T_c\)) increases significantly. Thus, even in the case that the glass transition is unobservable at usual heating rates (e.g. below 40 K/min), it comes to observable by rapid heating.\textsuperscript{12,13} As mentioned above, since raising the heating rates can significantly affect the crystallization behavior, it is a useful method to investigate the crystallization mechanism or primary crystallization. In this paper, we systematically investigate crystallization process and thermal stability of Zr\textsubscript{50}Cu\textsubscript{40}Al\textsubscript{10} BMG, exhibiting a broad supercooled liquid region, by using differential scanning calorimetry (DSC), X-ray diffraction, field-emission scanning electron microscopy (FE-SEM), X-ray diffraction method, long-range-diffusion, strain energy.

2. Experimental

Alloy ingots of Zr\textsubscript{30}Cu\textsubscript{40}Al\textsubscript{10}, Zr\textsubscript{70}Cu\textsubscript{30}, and Zr\textsubscript{70}Cu\textsubscript{20}Al\textsubscript{10} alloys were prepared by arc-melting a mixture of pure Zr, Cu, and Al metals in an argon atmosphere. In order to maintain a low-oxygen concentration in these alloys (especially for Zr\textsubscript{50}Cu\textsubscript{40}Al\textsubscript{10} bulk glass), Zr-crystal rods, specially prepared with oxygen concentration less than 0.05 at%, were used. Zr\textsubscript{50}Cu\textsubscript{40}Al\textsubscript{10} BMG was cast into block-shaped samples of \(3 \text{ mm} \times 3 \text{ mm} \times \text{length in dimensions. Ribbon-shaped Zr}\textsubscript{70}Cu\textsubscript{20}Al\textsubscript{10} and Zr\textsubscript{70}Cu\textsubscript{30} amorphous alloys were prepared by quenching from 1200°C, using a single roller melt-spinning technique in an argon atmosphere with a roller spinning rate at 4000 rpm. The densities of the samples

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measured by the Archimedean method are 6.837 g/cm³ for Zr₅₀Cu₄₀Al₁₀, 6.50 g/cm³ for Zr₇₀Cu₃₀Al₁₀ and 7.0 g/cm³ for Zr₇₀Cu₃₀, respectively. The density of the fully crystallized Zr₅₀Cu₄₀Al₁₀ is 6.879 g/cm³.

Differential scanning calorimetry (DSC) measurements were carried out using a standard commercial instrument (Perkin Elmer Diamond DSC) with about 5–10 mg samples. The characteristic temperatures, T_g and T_x, was measured during heating at a constant rate up to β = 500 K/min, which is the highest β attained in our present instrument. Crystalline phases in the samples after crystallization in the DSC apparatus were also investigated by X-ray diffraction with CuKα radiation. Structure observation, chemical composition analyses after crystallization were carried out by FE-SEM equipped with an energy dispersion X-ray (EDX) spectrometer.

3. Strain-energy Calculation

In order to evaluate the elastic strain energy supposing that the nucleation of the crystalline phase occurs in “an amorphous solid”, we employ the Eshelby inclusion theory. In that theory, it is assumed that an ellipsoidal crystalline particle is formed in the original solid glassy matrix, and the strain energy caused by the volume (density) change with crystallization was evaluated by using the following equation:

$$E_{\text{str}} = -\frac{1}{2} C_{ijkl} (S_{\text{lin}} e_{\text{lin}}^{\ast} - e_{\text{lin}}^{\ast}) x_{ijkl} \Omega_{x}$$

$$e_{\text{lin}}^{\ast} = \frac{\sqrt{1/\rho_{\text{Cryst}} - \sqrt{1/\rho_{\text{Amo}}}}}{\sqrt{1/\rho_{\text{Amo}}}}$$

where $C_{ijkl}$, $S_{\text{lin}}$, $\Omega_{x}$, $e_{\text{lin}}^{\ast}$ are the elastic constants, the Eshelby tensor, the volume of the inclusion, eigenstrain due to the volume mismatch (i.e., $e_{\text{lin}}^{\ast} = e^{\ast} \delta_{\text{lin}}$ with $\delta_{\text{lin}}$ the Kronecker delta), respectively, and $\rho$ is the mass density. In the present calculation, we assume that the elastic constants of the crystalline phases are the same as those of the amorphous phase, and we use the elastic constants of the Zr₅₀Cu₄₀Al₁₀ BMG measured with the resonant ultrasound spectroscopy (RUS) technique at room temperature.

4. Results and Discussion

4.1 Thermal stability of Zr₅₀Cu₄₀Al₁₀ BMG

The characteristic temperatures $T_g$ and $T_x$ of the Zr₅₀Cu₄₀Al₁₀ BMG, Zr₇₀Cu₃₀Al₁₀, and Zr₇₀Cu₃₀ amorphous alloys were measured with DSC from 5 K/min up to 500 K/min. Typical DSC profiles obtained for the Zr₅₀Cu₄₀Al₁₀ BMG at various heating rates are shown in Fig. 1(a). The glass transition temperature $T_g$ is defined as the onset temperature of the endothermic reaction, and the crystallization temperature $T_x$ is defined as the onset value of the exothermic peak. Figure 1(b) compared thermal stability of the three Zr-Cu-Al alloys; the $T_g$ and $T_x$ values for the three alloys are plotted against log $\beta$. It is found that these characteristic temperatures can be expressed in Lasocka’s empirical form:

$$T_{x,g} = a_{x,g} + b_{x,g} \log \beta$$

Ichitsubo et al. reported that the intersection point of the two fitted lines of $T_g$ and $T_x$ in the Lasocka plot, $\beta_c$, is one of indices for evaluating the thermal stability of BMGs. In general, the maximum diameter is frequently used for evaluation of GFA, which is an index of stability during a cooling process. In contrast, the parameter $\beta_c$ used here is an index for the thermal stability of the glassy solid during a heating process. A more stable BMG tends to exhibit a smaller value of $\beta_c$. Table 1 summarizes the $\beta_c$ values of several BMGs obtained from the present experiment and our previous experiments. The value of $\beta_c$ of the Zr₅₀Cu₄₀Al₁₀ BMG is found to be quite small and comparable to that of the Pd₄₂.₅Ni₇.₅Cu₃₀P₂₀ BMG which is well known as the highest glass former among metallic glasses. This indicates that Zr₅₀Cu₄₀Al₁₀ can also be regarded as one of the most stable BMGs. It is interesting to note that $\beta_c$ of the Zr₇₀Cu₃₀Al₁₀ glass is significantly...
smaller than that of the binary Zr$_{70}$Cu$_{30}$ amorphous alloy, despite the fact that $\Delta T_c (\equiv T_c - T_g)$ of the two glasses are not largely different. Thus, it is found that the three glasses show significant differences in terms of the thermal stability, although they are composed of the same constituent elements.

According to the structural analyses by X-ray diffraction for Zr$_{70}$Cu$_{30}$Al$_{10}$ and Zr$_{70}$Cu$_{30}$ amorphous alloys, the total coordination numbers around each constituent element are approximately 10. In contrast, the extremely high total coordination numbers exceeding 11 around all constituent elements in Zr$_{50}$Cu$_{40}$Al$_{10}$ are only realized in a homogeneously close-packed glassy structure. Additionally, average number densities of atoms ($\rho_0$) were calculated by

$$\rho_0 = \frac{N_A}{M} \rho_S, \quad \bar{M} = M_{Zr}X_{Zr} + M_{Cu}X_{Cu} + M_{Al}X_{Al},$$

where $N_A$, $\rho_S$, and $\bar{M}$ are the Avogadro number, sample mass density, and the average molecular mass, which is calculated using the atomic mass numbers of each constituent element ($M_{Zr} = 91.22, M_{Cu} = 63.55, M_{Al} = 26.98$) and the compositions ($X_{Zr}, X_{Cu}, X_{Al}$). The $\rho_0$ values of respective glasses are Zr$_{50}$Cu$_{40}$Al$_{10}$: 0.0537 (1/Å$^3$), Zr$_{70}$Cu$_{30}$Al$_{10}$: 0.0494 (1/Å$^3$), Zr$_{20}$Cu$_{50}$: 0.0508 (1/Å$^3$). The relatively high $\rho_0$ value of Zr$_{50}$Cu$_{40}$Al$_{10}$ BMG is also an indication of the homogeneously close-packed glassy structure.

Next, we discuss the crystallization behaviors of the respective glasses. Figure 2 shows the X-ray diffraction profiles of the samples after heated just above the crystallization temperature. Each sample was heated in the DSC profiles of the samples after heated just above the crystalline glasses. Figure 2 shows the X-ray diffraction profiles of the samples after heated just above the crystallization temperature. Each sample was heated in the DSC profiles of the samples after heated just above the crystalline glasses. As is found from Fig. 2, the crystallization process is rather simple in both cases of Zr$_{70}$Cu$_{30}$Al$_{10}$ and Zr$_{70}$Cu$_{30}$ and the main crystalline phase is Zr$_2$Cu with C11b structure. By contrast, Zr$_{50}$Cu$_{40}$Al$_{10}$ shows a complicated crystallization process, and the resulted crystalline phases are plural; apart from the principal Zr$_2$Cu crystalline phase, AlCu$_2$Zr (one of equilibrium phases in Zr$_{50}$Cu$_{40}$Al$_{10}$) and B2-type ZrCu (equilibrium phase higher than 715°C$^{21,22}$) are also formed. Especially, the latter phase, B2-type ZrCu, is a non-equilibrium phase, and this singular formation is deeply discussed in the later section. For such a complicated crystallization process including phase separation, long-range diffusion of atoms is generally required. Thus, it is considered that long-range diffusion of atoms plays an important role of retarding/preventing the crystallization in the Zr$_{50}$Cu$_{40}$Al$_{10}$ BMG.

4.2 Formation of B2-type ZrCu phase in Zr$_{50}$Cu$_{40}$Al$_{10}$

In order to clarify the complicated crystallization of the Zr$_{50}$Cu$_{40}$Al$_{10}$ BMG, we focus on the heating-rate dependence of its crystallization behavior. As is seen in Fig. 1(a), the supercooled liquid region ($\Delta T_c$) of Zr$_{50}$Cu$_{40}$Al$_{10}$ is as broad as 100 K, even when $\beta$ is less than 40 K/min. It is seen that $\Delta T_c$ becomes broader at a higher $\beta$, because the increase in $T_c$ with increasing $\beta$ is much steeper than that in $T_g$. Since the crystallization of the amorphous phase generally requires atomic diffusion, when $\beta$ is higher it is more difficult to gain sufficient atomic diffusion distance for crystallization.

Figure 3(a) shows XRD profiles obtained for the Zr$_{50}$Cu$_{40}$Al$_{10}$ glass heated to 700°C at respective heating rates. (b) The intensity ratio of two diffraction peaks (110 from B2 CuZr and 103 from C11b Zr$_2$Cu) after crystallized at various rates.
becomes higher. This is summarized in Fig. 3(b), which shows the intensity ratio of two X-ray diffraction peaks (110 from ZrCu and 103 from Zr₅₀Cu₄₀) after crystallized at various rates. As far as we know, this is the first time to report that B₂-type ZrCu is formed in the Zr-Cu-Al metallic glass system.

In the case of a higher β, the spent time till reaching a certain temperature is shortened, which means that a higher β reduces the annealing time at sufficiently high temperatures where atomic diffusion appreciably takes place, that is, the long-range diffusion of atoms enough for crystallization cannot be taken place. In other words, the present result strongly means that, when a glass phase has to be crystallized without long-range diffusion of atoms, it tends to transform into a crystalline phase as the primary phase, that possesses a similar composition and/or structure to those of the parent phase. Consequently, Zr₅₀Cu₄₀ possesses a considerable high thermal stability among metallic glasses.

Table 2: Interatomic distances and coordination number in Zr₅₀Cu₄₀Al₁₀ and B₂-ZrCu.

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<th>Zr₅₀Cu₄₀Al₁₀ [r (nm)]</th>
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<td>Zr-Zr</td>
<td>0.323</td>
<td>0.322</td>
<td>5.8</td>
<td>6</td>
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<tr>
<td>Zr-Cu</td>
<td>0.286</td>
<td>0.279</td>
<td>4.7</td>
<td>8</td>
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<tr>
<td>Zr-Al</td>
<td>0.305</td>
<td>—</td>
<td>—</td>
<td>—</td>
</tr>
<tr>
<td>Total</td>
<td>11.8</td>
<td>—</td>
<td>14</td>
<td></td>
</tr>
<tr>
<td>Cu-Zr</td>
<td>0.286</td>
<td>0.279</td>
<td>5.9</td>
<td>8</td>
</tr>
<tr>
<td>Cu-Cu</td>
<td>0.265</td>
<td>0.322</td>
<td>4.0</td>
<td>6</td>
</tr>
<tr>
<td>Cu-Al</td>
<td>0.269</td>
<td>—</td>
<td>0.9</td>
<td>—</td>
</tr>
<tr>
<td>Total</td>
<td>10.7</td>
<td>—</td>
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Table 2 compares the local structure of the Zr₅₀Cu₄₀Al₁₀ BMG determined by X-ray diffraction⁹ and the structure of B₂-type ZrCu.²¹,²² The Zr-Cu distance in Zr₅₀Cu₄₀Al₁₀ is 0.286 nm, being very close to the distance (0.279 nm) of the nearest neighbor Zr-Cu pair in B₂ ZrCu. In addition, the Zr-Zr distance (0.323 nm) in Zr₅₀Cu₄₀Al₁₀ is also close to the distance (0.323 nm) of the next nearest neighbor Zr-Zr pair in B₂-ZrCu. Thus, it may be possible to say that the local structures of the original glassy matrix and the ordered B₂-type ZrCu have certain similarities in terms of the interatomic distances and atomic coordination. Here, we encounter a new problem, that is, we have to consider why the B₂ phase is not easily formed in the usual crystallization behavior of the Zr₅₀Cu₄₀Al₁₀ BMG.

As is mentioned above, the ordered B₂-type ZrCu is a high-temperature equilibrium phase (higher than 715°C)²¹,²² in the binary alloy); hence, around about 650°C (this was the maximum temperature in the present DSC experiments), the B₂-type ZrCu phase is a non-equilibrium phase. Thus, as is seen in Fig. 4, it is naturally expected that the driving force for the transformation into the ordered B₂-type ZrCu phase should be not so large, compared to that for the transformation into the thermodynamically equilibrium phases such as C₁₁₅-type Zr₂Cu and the others. Repeatedly to say, long-range diffusion is required for the latter crystallization process including the phase separation. Consequently, Zr₅₀Cu₄₀Al₁₀ possesses a considerable high thermal stability among metallic glasses.

Fig. 4 Schematic diagram showing the decrease of the chemical-free-energy accompanied by the crystallization process. The solid (red) arrows indicate the crystallization process driven by the short-range diffusion, and the hatched (blue) arrows indicate the long-range-diffusion process.

4.3 Effect of elastic strain energy

It is well known that the elastic strain contributes to retarding or prohibiting the nucleation in a mother phase. Therefore, for discussing the presence or absence of stable glass-to-liquid transition of metallic glasses, it is meaningful to evaluate the elastic strain energy associated with the crystallization. For the purpose of the present calculation, we consider that the nucleation of the crystalline phase occurs in the solid amorphous matrix. Table 3 summarizes the results of the calculated elastic strain energy, $E_{\text{str}}$, caused by crystallization in several glasses. The elastic strain energy was calculated using the expected mass density, $\rho_{\text{cryst}}$, of the crystal, which is estimated under the condition that the single phase with the B₂ or C₁₁₅ structure of the lattice parameters measured by X-ray diffraction contains the constituent atoms in accordance with the atomic compositions of the mother glass. The crystallization heat, $\Delta H$, was obtained from the DSC profile at $\beta = 40$ K/min. The elastic constants reported by Zhang et al.³ were used in the present calculation.

Since the crystallization heat is the enthalpy difference between the amorphous state and crystallized state, we may regard that the driving force of crystallization is approximately proportional to the magnitude of $\Delta H$. On the other hand, the strain energy caused by the volume (density) change accompanied by the crystallization should be a positive effect on preventing the crystallization. Therefore, it is reasonable to claim that the stable glass-to-liquid transition will appear when the magnitude of $E_{\text{str}}/\Delta H$ cannot be neglected, because the nucleation of crystals becomes easier in the supercooled liquid state that has softer elastic constants. As is seen in Table 3, the $E_{\text{str}}/\Delta H$ value for the Zr₇₀Cu₃₀ amorphous alloy is considerably smaller than the values of Zr₅₀Cu₄₀Al₁₀ and Zr₅₀Cu₃₀Al₁₀ metallic glasses. Thus, a large amount of the elastic strain energy, i.e., large value of $E_{\text{str}}/\Delta H$, contributes to the prevention of crystallization, which leads to the excellent thermal stability of BMGs.

Here we especially focused on the strain energy associated with the volume change by crystallizing to sound out
any other origin of the formation of B2 ZrCu as the primary phase instead of an equilibrium C11$_b$ Zr$_2$Cu phase. From the FESEM-EDX analysis, we have found that the compositional deviation of Al from the composition of the mother glass is less than 2 at% even after annealing at a sufficiently high temperature for long duration (600°C, 22 h). Thus, it appears to be reasonable to assume that the C11$_b$-type Zr$_2$Cu crystalline phase is formed with the same composition of the Zr$_{50}$Cu$_{40}$Al$_{10}$ mother glass. Since the composition of Zr$_2$Cu is obviously different from the original glass, a large quantity of Cu and Al atoms have to occupy the Zr sites in Zr$_2$Cu instead of Zr atoms on the present assumption. Thus, the density of C11$_b$-type Zr$_2$Cu is expected to decrease, being far from the mass density of the glassy solid, so that the value of $E_{str}/\Delta H$ inevitably amounts to about 30%, if such a crystalline phase is supposed to be formed in the glass matrix. Thus, we can understand from the viewpoints of the strain energy that the formation of the equilibrium Zr$_2$Cu phase in the glassy solid is quite unlikely. Therefore, in order for the Zr$_{50}$Cu$_{40}$Al$_{10}$ glass to be crystallized easily, first the glass matrix has to become softer by transforming into a supercooled liquid, and then long-range diffusion of Cu and Al atoms are required for the complicated crystallization process including a phase separation.

On the other hand, although B2-type ZrCu is a high-temperature equilibrium phase, the strain energy is not so large for the crystallization to B2-type ZrCu, and furthermore long-range diffusion of Al atoms is not required because the composition is similar to that of the original glass matrix, resulting in the facile crystallization toward B2-type ZrCu.

5. Conclusion

We have investigated the complicated crystallization behavior and origin of the high thermal stability of the Zr$_{50}$Cu$_{40}$Al$_{10}$ BMG, by comparing to Zr$_{70}$Cu$_{30}$ and Zr$_{90}$Cu$_{10}$ amorphous alloys. The salient results are summarized as follows:

1. The Zr$_{50}$Cu$_{40}$Al$_{10}$ BMG has a homogeneously close-packed glassy structure, compared to the other amorphous alloys.

2. It is found that the unusual B2-type ZrCu crystalline phase is mainly formed when the glass is rapidly heated up. The reason of formation of the B2 phase is that the long-range diffusion cannot take place sufficiently in the case of high heating rates and, hence, the B2-type ZrCu phase possessing a similar composition to that of the glass matrix is favored to be formed. From this result, it is considered that the B2 phase is probably a primary phase in the usual crystallization.

3. The B2-type ZrCu is a high-temperature phase and not an equilibrium phase around $T_g$ and $T_s$ of the glass; therefore, the driving force of the crystallization toward the B2 phase is considered to be lower than that in the case of the crystallization to the equilibrium phases (e.g., Zr$_2$Cu), and furthermore long-range diffusion is inevitable required for the latter crystallization process including the phase separation. This is an origin of the high thermal stability and main reason for retarding the crystallization in the Zr$_{50}$Cu$_{40}$Al$_{10}$ BMG.

4. It is considered that the elastic strain plays a key role in the presence of the stable glass-to-liquid transition of Zr-M and Zr-Al-M (M=Cu, Ni) metallic glasses. It is found that the ratio of the strain energy to the crystallization heat ($E_{str}/\Delta H$) of the Zr$_{50}$Cu$_{40}$Al$_{10}$ BMG is the highest between the three glasses. Moreover, the value of $E_{str}/\Delta H$ of Zr$_{70}$Cu$_{30}$Al$_{10}$ exceeds that of Zr$_{70}$Cu$_{30}$. This indicates that the value of $E_{str}/\Delta H$ reflects the thermal stability of the metallic glasses upon heating treatments.

Acknowledgements

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