

Studies on the Electrocatalytic Reduction of Hydrogen Peroxide on a Glassy Carbon Electrode Modified With a Ruthenium Oxide Hexacyanoferrate Film

Roselyn C. Peña¹, Juan C. M. Gamboa¹, Mauro Bertotti¹ and Thiago R. L. C. Paixão^{1,2,*}

¹ Instituto de Química - Universidade de São Paulo, São Paulo, SP, Brazil 05508-900

² Centro de Ciências Naturais e Humanas, Universidade Federal do ABC, Santo André, SP, Brazil 09210-170

*E-mail: trlcp@iq.usp.br

Received: 5 August 2010 / *Accepted:* 15 January 2011 / *Published:* 1 February 2011

The electrocatalytic reduction of hydrogen peroxide on a glassy carbon (GC) electrode modified with a ruthenium oxide hexacyanoferrate (RuOHCF) was investigated using rotating disc electrode (RDE) voltammetry aiming to improve the performance of the sensor for hydrogen peroxide detection. The influence of parameters such as rotation speed, film thickness and hydrogen peroxide concentration indicated that the rate of the cross-chemical reaction between Ru(II) centres immobilized into the film and hydrogen peroxide controls the overall process. The kinetic regime could be classified as LSk mechanism, according to the diagnostic table proposed by Albery and Hillman, and the kinetic constant of the mediated process was found to be $706 \text{ mol}^{-1} \text{ cm}^3 \text{ s}^{-1}$. In the LSk case the reaction layer is located at a finite layer close to the modifier layer/solution interface

Keywords: Albery-Hillman's diagnostic table, modified electrode, rotating disc electrode, electrocatalysis, hydrogen peroxide.

1. INTRODUCTION

Amperometric sensors for hydrogen peroxide present several applications, such as in environmental, pharmaceutical, medicine, food samples and as disinfectant and bleaching agent in various industrial and household products [1-3]. Their advantages are based on low detection limits, large dynamic concentration range, good selectivity, rapid response time and inherent miniaturization and portability. However, these features depend on the use of mediators capable of recognizing the analyte in complex samples. Several chemical compounds have been attached onto the electrode to minimize the energy required in the electrochemical reduction of hydrogen peroxide [4-12]. Recently,

we have reported the development of a glassy carbon electrode coated with an electrodeposited film of ruthenium(III) oxide hexacyanoferrate (RuOHCF) for the determination of 2'-deoxyguanosine [13], ascorbate [14, 15] and hydrogen peroxide [16]. At surfaces containing this chemical modifier, hydrogen peroxide is electrocatalytically reduced at 0.0 V, allowing amperometric determinations with a low limit of detection ($1.7 \mu\text{mol L}^{-1}$) [17] and good selectivity. To our knowledge, until now, there is no detailed study on the mechanism of the electrocatalytic reduction of H_2O_2 on RuOHCF films. Hence, in the present work the kinetics of the reaction involving hydrogen peroxide and immobilized Ru (II) centers was investigated by using rotating disc electrode (RDE) voltammetry. Data were analyzed with the model developed by Albery and Hillman [18, 19], which allows a full characterization of the electrocatalytic process based on the limiting current dependence upon rotation speed, film thickness and substrate concentration.

2. EXPERIMENTAL SECTION

2.1. Chemicals, materials and samples

All solid reagents were of analytical grade and were used without further purification. The solutions were prepared by dissolving the reagents in deionized water processed through a water purification system (Nanopure Infinity, Barnstead). Potassium ferricyanide, potassium chloride, potassium permanganate, hydrogen peroxide, sulfuric acid and ferric chloride were obtained from Merck (Darmstadt, Germany) and ruthenium (III) chloride was obtained from Alfa Aesar (Massachusetts, USA). Hydrogen peroxide solutions were daily prepared from a 30% m/m stock solution and standardized as reported in the literature [20].

2.2. Electrodes and instrumentation

An Autolab PGSTAT 30 (Eco Chemie) bipotentiostat with data acquisition software made available by the manufacturer (GPES 4.8 version) was used for electrochemical measurements. A homemade Ag/AgCl (saturated KCl) and a platinum wire were used as reference and counter electrodes, respectively. Voltammetry with a glassy carbon rotating disc electrode (RDE) was carried out using an analytical rotator (AFMSRX, Pine Instrument Company) connected to the potentiostat. Atomic force microscope images were obtained with a 5500 Atomic Force Microscope (AFM) (N9410S). Experiments were carried out using a 1 cm^2 glassy carbon plate (Alfa Aesar, Massachusetts, USA).

2.3. Modification of the Electrode Surface

The surface of the GC electrode was polished with alumina suspension ($1 \mu\text{m}$, Alfa Aesar, Massachusetts, USA) on a microcloth polishing pad, rinsed with water and sonicated for five minutes in distilled water. The electrodeposition of the ruthenium oxide hexacyanoferrate film onto the surface

of the glassy carbon electrode was performed by repetitive sweeping at the potential limits of -0.5 and 1.3 V at 100 mV s^{-1} in a solution containing $0.5 \text{ mol L}^{-1} \text{ KCl} + 0.05 \text{ mol L}^{-1} \text{ HCl} + 1 \text{ mmol L}^{-1} \text{ K}_3\text{Fe}(\text{CN})_6 + 1 \text{ mmol L}^{-1} \text{ RuCl}_3$. The stability of the modified electrode was verified by recording cyclic voltammograms with the modified electrode in a supporting electrolyte solution containing $0.5 \text{ mol L}^{-1} \text{ KCl} + 0.05 \text{ mol L}^{-1} \text{ HCl}$ at the same potential limits.

3. RESULTS AND DISCUSSION

3.1. Electrochemical behavior of the RuOHCF modified electrode

AFM was used to confirm the electrodeposition of the RuOHCF film. Fig.1 depicts 3D views of AFM images obtained for a freshly polished bare glassy carbon electrode before and after the film deposition.

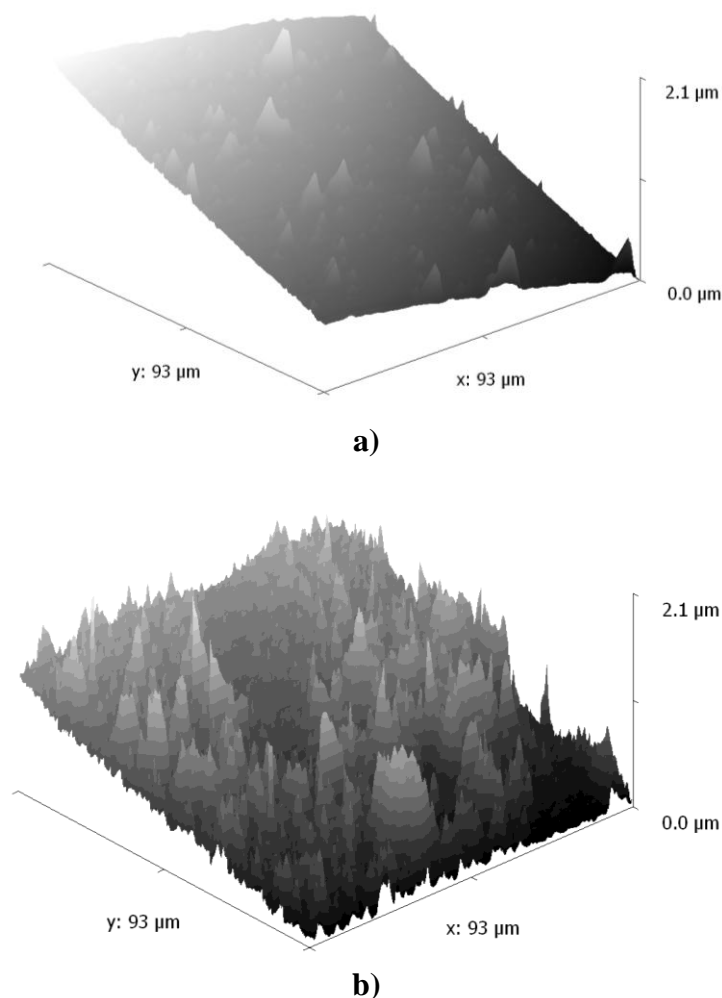


Figure 1. AFM-images of the glassy carbon surface before (A) and after (B) electrodeposition of a RuOHCF film ($\Gamma = 9.16 \times 10^{-9} \text{ mol cm}^{-2}$).

At the conditions described in the Experimental section for the film deposition [13, 21], an extremely rough surface is obtained as a consequence of the growth of an insoluble material with a heterogeneous morphology, and the surface area is largely increased (average maximum height of the roughness 0.052 and 0.485 μm , respectively, for a bare and a modified electrode).

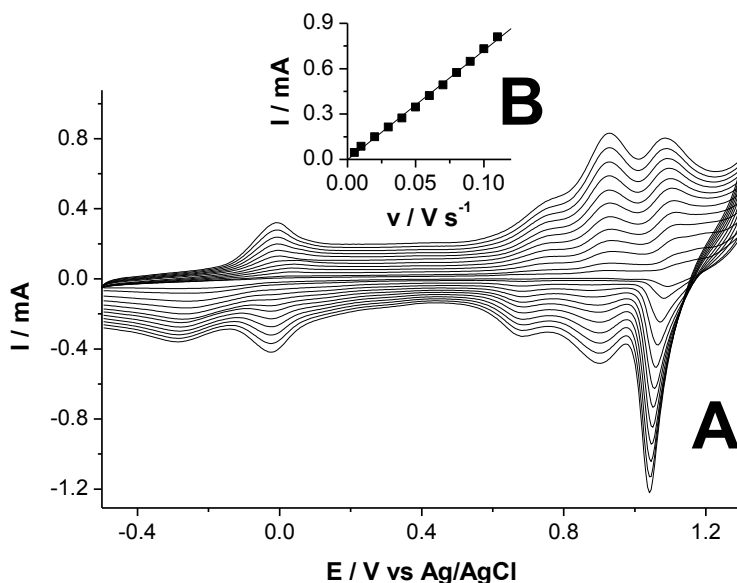
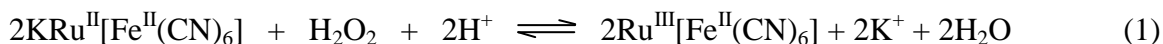


Figure 2. Cyclic voltammograms recorded in a solution containing 0.5 mol L⁻¹ KCl and 0.05 mol L⁻¹ HCl using a RuOHCF modified electrode ($\Gamma = 9.16 \times 10^{-9}$ mol cm⁻²) at different scan rates (A). Linear relationship between peak current values (measured at 0.93 V) and scan rate (B).

The effect of the scan rate on the electrochemical behavior of the RuOHCF modified electrode was studied as shown in Fig. 2. A linear correlation between current (measured at 0.93 V) corresponding to the Ru(III)/Ru(IV) couple and scan rate indicates that the flow of electrons inside the film is fast. Fig 2 B shows a linear relationship between peak current and scan rate, contrary to what is expected when the process is controlled by diffusion of an electroactive species in solution. Such behavior is typical for modified electrodes that exhibit reversible redox process [22].

3.2. Electrocatalytic reduction of hydrogen peroxide

The electrocatalytic activity of the modified electrode towards the cathodic reduction of hydrogen peroxide is presented in Fig. 3. After the addition of hydrogen peroxide, a significant current enhancement is seen at around 0.0 V with a simultaneous decrease in the corresponding anodic peak. This observation is a clear evidence of electrocatalysis, attributable to the reaction of electrogenerated Ru(II) centers with diffusing hydrogen peroxide according to the following equation:



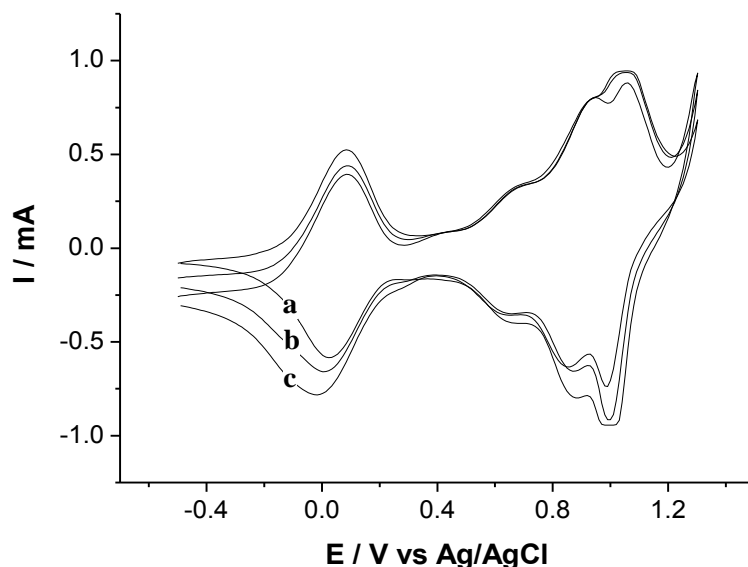


Figure 3. Cyclic voltammograms recorded in a solution containing $0.5 \text{ mol L}^{-1} \text{ KCl} + 0.05 \text{ mol L}^{-1} \text{ HCl}$ before (a) and after (b,c) addition of hydrogen peroxide (final concentration 2.45 (b) and 4.82 (c) mmol L^{-1}) by using a RuOHCF modified electrode ($\Gamma = 9.16 \times 10^{-9} \text{ mol cm}^{-2}$). Scan rate = 25 mV s^{-1}

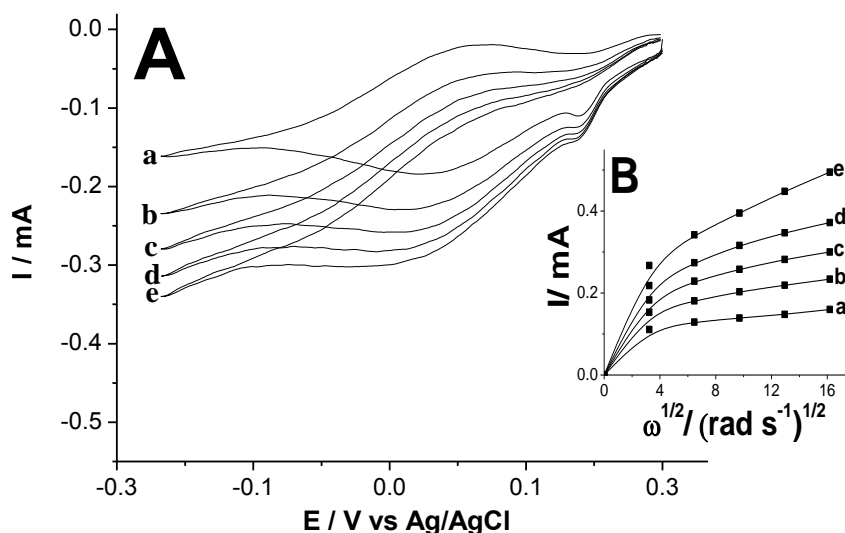


Figure 4. Cyclic voltammograms (A) recorded in a solution containing $2.45 \text{ mmol L}^{-1} \text{ H}_2\text{O}_2 + 0.5 \text{ mol L}^{-1} \text{ KCl} + 0.05 \text{ mol L}^{-1} \text{ HCl}$ using a RuOHCF modified electrode ($\Gamma = 9.16 \times 10^{-9} \text{ mol cm}^{-2}$) at different rotation speeds: 100 rpm (a), 400 rpm (b), 900 rpm (c), 1600 rpm (d) and 2500 rpm (e). Scan rate: 20 mV s^{-1} . Levich plots (B) for the reduction of hydrogen peroxide at different concentrations: 1.00 (a), 1.82 (b), 2.45 (c), 3.43 (d) and 4.82 mmol L^{-1} (e) using a RuOHCF modified electrode.

In order to investigate in more details the electrocatalytic process, modified electrodes were used in rotating disc electrode voltammetric experiments. Data were obtained with different rotation speeds, film thickness and hydrogen peroxide concentration to obtain information on the possible rate-

determining step as well as on the location of the reaction zone. Fig. 4A shows voltammograms recorded at different rotation speeds with a RuOHCF modified electrode ($\Gamma = 9.16 \times 10^{-9} \text{ mol cm}^{-2}$) in supporting electrolyte solution containing hydrogen peroxide at a fixed concentration. Enhanced limiting currents in voltammograms recorded at increased mass transport conditions are observed. Similar RDE voltammetric experiments were carried out by varying the concentration of hydrogen peroxide. The Levich plots shown in Fig. 4B demonstrate that at relatively high rotation rates the electrode response to hydrogen peroxide is no longer only limited by diffusion, but also by the kinetics of the reaction between catalytic sites of the film and the substrate.

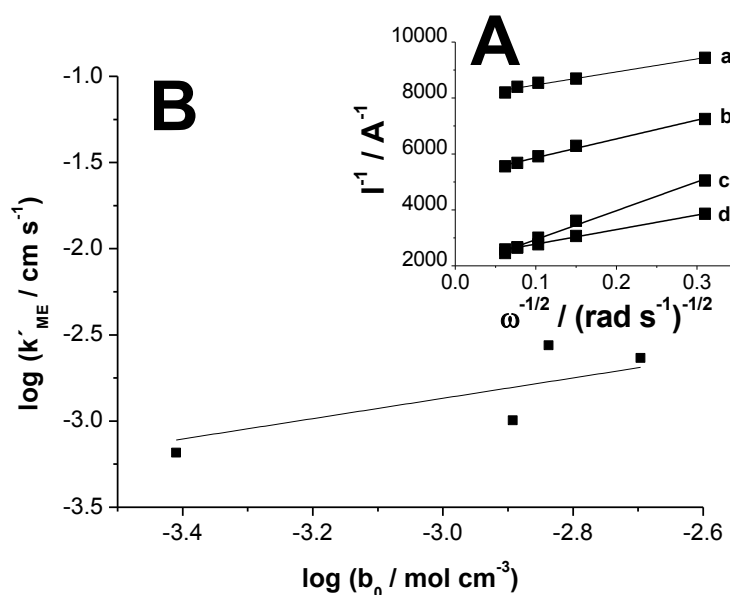


Figure 5. Koutecký-Levich (A) plots for reduction of hydrogen peroxide in a solution containing $1.00 \text{ mmol L}^{-1} \text{ H}_2\text{O}_2 + 0.5 \text{ mol L}^{-1} \text{ KCl} + 0.05 \text{ mol L}^{-1} \text{ HCl}$ using a RuOHCF modified electrode with different Γ values: $1.54 \times 10^{-9} \text{ mol cm}^{-2}$ (a), $9.16 \times 10^{-9} \text{ mol cm}^{-2}$ (b), $1.04 \times 10^{-8} \text{ mol cm}^{-2}$ (c) and $1.44 \times 10^{-8} \text{ mol cm}^{-2}$ (d). Dependence of $\log k'_{ME}$ vs $\log b_0$ (B).

Information on the limiting dynamic steps was achieved by minimizing the influence of mass-transport of hydrogen peroxide to the film surface. This is typically done using the limiting current at an infinite rotation rate via a Koutecký-Levich plot. At this condition, charge transport, substrate diffusion within the film and the cross-chemical reaction may determine the value of $1/I$ at $\omega^{-1/2} = 0$. Accordingly, Fig. 5A shows Koutecký-Levich plots for four RuOHCF modified electrodes containing different film amounts (from 1.54×10^{-9} to $1.04 \times 10^{-8} \text{ mol cm}^{-2}$). The kinetic limitation of the electrode process is confirmed by considering the intercept values at $\omega^{-1/2} = 0$ of all linear plots, which are positive. The intercept corresponds to infinite rotation rate with negligible concentration polarization in the electrolyte solution and can be used to get information on the mediated reaction. As intercept values are inversely proportional to the concentration of the diffusing substrate, current is not limited by charge diffusion within the film layer and the reaction rate is controlled by transport of hydrogen peroxide within the film and/or by the cross-exchange reaction. For such an electrodic system the

characteristic current may be determined by the Koutecký-Levich analysis by using the relationship:

$$\frac{1}{I_{\text{lim}}} = \frac{1}{nFAc_s} \left[\frac{1}{Lev\omega^{1/2}} + \frac{1}{k'_{ME}} \right] \quad (2)$$

$$Lev = 0.62D^{2/3}\nu^{-1/6} \quad (3)$$

where A is the area of the electrode, c_s is the substrate concentration at the bulk solution, ω is the rotation speed in rad s^{-1} and k'_{ME} is the heterogeneous rate constant of the electrochemical process. Lev is the Levich constant and is given by equation 3 [23, 24], where D is the diffusion coefficient of the substrate in the electrolyte and ν is the kinematic viscosity. k'_{ME} was determined using the values of the intercept at $\omega^{-1/2} = 0$ (Figure 5A), $n = 2$ for the reduction process of hydrogen peroxide and the Equation 2 for each experimental condition.

By taking into account the calculated k'_{ME} values, it is possible to apply the diagnostic procedure [24] reported by Alberly and Hillman [18, 19, 24, 25] to elucidate the kinetics of the overall process. As limiting current is dependent on rotation rate and Koutecký-Levich plots are linear (Fig. 4 and 5), it is necessary to calculate the value of the Levich constant, Lev, from the gradient of the linear Koutecký-Levich plots (Fig. 5A) and compare the result with the one obtained by using a bare electrode. The mean value of the Lev constant was found to be $(8 \pm 2) \times 10^{-4} \text{ cm s}^{-1/2}$, in good agreement with the value obtained for the Lev constant ($8.9 \times 10^{-4} \text{ cm s}^{-1/2}$), which was calculated by using the equation 3 (diffusion coefficient of hydrogen peroxide in solution = $1.71 \times 10^{-5} \text{ cm}^2 \text{ s}^{-1}$ [26, 27] and $\nu = 0.01 \text{ cm}^2 \text{ s}^{-1}$). Since both Lev values agreed, there were only four possible cases to consider: Sk', LSk, Lk and LEt_y. In order to decide which case applies, the order of k'_{ME} versus b_0 (concentration of electrocatalytic sites into the film, calculated as described in a previous work [21]) was analyzed in a log-log plot (Figure 5B). The slope of the straight line was found to be 0.59 (~0.5), which suggests an LSk case taking into account the diagnostic scheme proposed by Alberly and Hillman [24]. According to this model, the reaction takes place at a finite layer close to the film/solution interface.

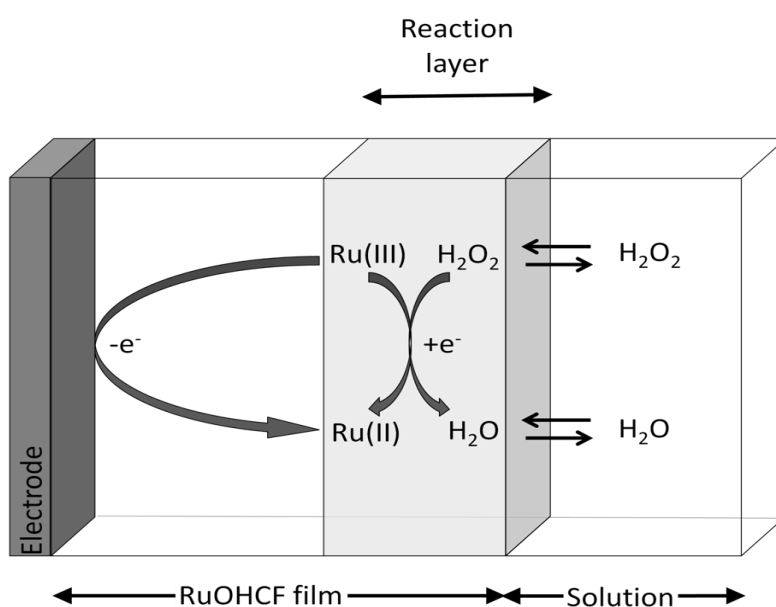
The rate constant of the mediated process (k) in the LSk case can be determined according to the following equation:

$$k'_{ME} = \kappa (k b_0 D_Y)^{1/2} \quad (4)$$

where b_0 is the concentration of electrocatalytic sites into the film ($1.09 \times 10^{-3} \text{ mol cm}^{-3}$) [21], D_Y is diffusion coefficient of hydrogen peroxide in the modifier layer (an approximation to the value obtained in biofilms at 25 °C was used ($7 \times 10^{-6} \text{ cm}^2 \text{ s}^{-1}$) [28]), and κ is the partition coefficient of the substrate between the film and the solution (usually use a value of 1). With the k'_{ME} value from Figure 5A (curve d) ($2.3 \times 10^{-3} \text{ cm s}^{-1}$) and the equation 5, the rate constant (k) was calculated as 693 mol^{-1}

$\text{cm}^3 \text{s}^{-1}$ (for a surface excess of $1.44 \times 10^{-8} \text{ mol cm}^2$), which indicates a low electrocatalytic efficiency when compared with other metal-hexacyanoferrates such as SnHCF [29] and NdHCF [30].

The reaction zone is close to the film/solution interface for an electroc process following the LSk case. Accordingly, it is expected that the flow of charge ($D_{\text{ct}}b_0/L$) is much higher than the flow of species resulting from the kinetic step ($k\Gamma_{\text{c}_s}$). By calculating these two parameters using D_{ct} , L and b_0 values previously reported [21], $D_{\text{ct}}b_0/L$ and $k\Gamma_{\text{c}_s}$ were found to be 5×10^{-9} and $1 \times 10^{-11} \text{ mol cm}^{-2} \text{ s}^{-1}$, respectively. As $D_{\text{ct}}b_0/L \gg k\Gamma_{\text{c}_s}$, the assumption that the reaction is located in the film/solution interface is confirmed, in agreement with the LSk mechanism. Considering the results obtained in this work, a possible mechanistic proposal for the process involving the electrocatalytic reduction of hydrogen peroxide at RuOHCF modified platforms is illustrated in the Scheme 1.



Scheme 1.

4. CONCLUSIONS

RDE studies indicated that the rate of cross-chemical reaction between Ru(II) centers immobilized into the film and hydrogen peroxide controls the limiting current and the value of the rate constant was found to be $706 \text{ mol}^{-1} \text{ cm}^3 \text{ s}^{-1}$. The kinetic regime may be classified as an LSk mechanism, according to the Albery and Hillman model. In the LSk case the reaction takes place at a finite layer close to the film/solution interface, indicating that diffusion into the film does not consist of one of the limiting steps of the overall process.

5. ACKNOWLEDGEMENTS

The authors are thankful to FAPESP (Fundação de Amparo à Pesquisa do Estado de São Paulo,

Grant Number: 09/07859-1), CAPES (Coordenação de Aperfeiçoamento de Pessoal de Nível Superior) and CNPq (Conselho Nacional de Desenvolvimento Científico e Tecnológico) for the financial support.

Appendix

Acronyms used in the manuscript:

b_0 = Concentration of catalytic centers at the modify layer, $mol\ cm^{-3}$

A = Geometric area, cm^2

L = Coating thickness, cm

κ = Partition coefficient of the substrate between film and solution, *dimensionless*

D_y = Diffusion coefficient of the substrate into the film, $cm^2\ s^{-1}$

D_{ct} = Diffusion coefficient of charge transport in the film, $cm^2\ s^{-1}$

D = Diffusion coefficient of the substrate in the solution, $cm^2\ s^{-1}$

The notation used to describe the Albery-Hillman models defines the location of the reaction into the film: L, throughout the layer; LE, near the electrode interface; LS, near the electrolyte interface; LRZ, reaction zone in the middle of the layer and S, at the electrolyte layer. The subscribed notation indicates whether kinetics (k) or mass-transport (t_y or t_e (transport of reactant or electrons, respectively) controls the overall process.

References

1. I.L. de Mattos, K.A. Shiraishi, A.D. Braz, and J.R. Fernandes, *Quim. Nova* 26 (2003) 373.
2. C.J. Tredwin, S. Naik, N.J. Lewis, and C. Scully, *Br. Dent. J.* 200 (2006) 371.
3. R.S. Freire, R. Pelegrini, L.T. Kubota, N. Duran, and P. Peralta-Zamora, *Quim. Nova* 23 (2000) 504.
4. M.S. Lin, T.F. Tseng, and W.C. Shih, *Analyst* 123 (1998) 159.
5. Q.J. Chi and S.J. Dong, *Anal. Chim. Acta* 310 (1995) 429.
6. A.A. Karyakin, E.E. Karyakina, and L. Gorton, *Talanta* 43 (1996) 1597.
7. E. Bustos and L.A. Godínez, *Int. J. Electrochem. Sci.*, 6 (2011) 1.
8. A.A. Karyakin, E.E. Karyakina, and L. Gorton, *Electrochem. Commun.* 1 (1999) 78.
9. A.A. Karyakin, E.E. Karyakina, and L. Gorton, *Anal. Chem.* 72 (2000) 1720.
10. K.S. Tseng, L.C. Chen, and K.C. Ho, *Sens. Actuator B-Chem.* 108 (2005) 738.
11. A.A. Karyakin, E.A. Puganova, I.A. Budashov, I.N. Kurochkin, E.E. Karyakina, V.A. Levchenko, V.N. Matveyenko, and S.D. Varfolomeyev, *Anal. Chem.* 76 (2004) 474.
12. F. Ricci, C. Goncalves, A. Amine, L. Gorton, G. Palleschi, and D. Moscone, *Electroanalysis* 15 (2003) 1204.
13. T.R.L.C. Paixão, C.C.M. Garcia, M.H.G. Medeiros, and M. Bertotti, *Anal. Chem.* 79 (2007) 5392.
14. T.R.L.C. Paixão, L.F. Barbosa, M.T. Carri, M.H.G. Medeiros, and M. Bertotti, *Analyst* 133 (2008) 1605.
15. T.R.L.C. Paixão and M. Bertotti, *Electrochem. Commun.* 10 (2008) 1180.
16. T.R.L.C. Paixão and M. Bertotti, *Electroanalysis* 20 (2008) 1671.
17. R.C. Peña, J.C.M. Gamboa, T.R.L.C. Paixão, and M. Bertotti, *Microchim. Acta* 166 (2009) 277.
18. W.J. Albery, M.J. Eddowes, H.A.O. Hill, and A.R. Hillman, *J. Am. Chem. Soc.* 103 (1981) 3904.

19. W.J. Albery and A.R. Hillman, *J. Electroanal. Chem.* 170 (1984) 27.
20. J. Mendham, R.C. Denney, J.D. Barnes, and M.J.K. Thomas, *Vogel's Quantitative Chemical Analysis*, Prentice Hall, London, 2000.
21. T.R.L.C. Paixão and M. Bertotti, *Electrochim. Acta* 52 (2007) 2181.
22. S.M. Chen, M.F. Lu, and K.C. Lin, *J. Electroanal. Chem.* 579 (2005) 163.
23. A.Ciszewski and G. Milczarek, *J. Electroanal. Chem.* 469 (1999) 18.
24. W.J. Albery, M.G. Boutelle, and A.R. Hillman, *J. Electroanal. Chem.* 182 (1985) 99.
25. M. Vilas-Boas, E.M. Pereira, C. Freire, and A.R. Hillman, *J. Electroanal. Chem.* 538 (2002) 47.
26. D.M.H. Kern, *J. Am. Chem. Soc.* 76 (1954) 4208.
27. M.V. Vazquez, S.R. Desanchez, E.J. Calvo, and D.J. Schiffrin, *J. Electroanal. Chem.* 374 (1994) 179.
28. Z.Q. Gao, Y.M. Zhang, M.L. Tian, and Z.F. Zhao, *J. Electroanal. Chem.* 358 (1993) 161.
29. H. Razmi and A. Taghvimi, *Int. J. Electrochem. Sci.* 5 (2010) 751.
30. Q.L. Sheng, H. Yu, and J.B. Zheng, *Electrochim. Acta* 52 (2007) 4506.