

## Note

### <sup>15</sup>N-labeled glyphosate synthesis and its practical effectiveness

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**ABSTRACT:** Glyphosate is the herbicide with the greatest participation in the world market, and it has been used intensively in agriculture for over 30 years, especially because of its relatively low cost and high effectiveness on different species. This study describes a synthesis method of glyphosate in the form of the isopropylamine salt enriched in the stable nitrogen isotope (<sup>15</sup>N). The <sup>15</sup>N-herbicide synthesis was performed using a phosphomethylation reaction with dialkyl phosphite and <sup>15</sup>N-glycine. The tests were conducted at a microscale and equimolar quantities. A 20% yield was obtained at the established conditions. To test the effectiveness of the synthesized glyphosate, an experiment was performed in a growth chamber with the weed *Lolium multiflorum* (ryegrass), which is mentioned to be effectively controlled as indicated in package inserts of isopropylamine salt-based products (glyphosate). The experimental design was completely randomized, with four replicates, and treatments consisted of isopropylamine salt sources at the commercial dose recommended for the product (720 g a.i. ha<sup>-1</sup>). The production of green and dry phytomass of the above-ground part and roots was evaluated 21 days after application (DAA), and the treatment toxicity was evaluated 7, 14, and 21 DAA. The isopropylamine salt sources did not differ, recognizing them as useful isotopic tracers.

**Key words:** *Lolium multiflorum*, isopropylamine salt, stable isotope

### Síntese do glifosato marcado com nitrogênio-15 e sua eficiência prática

**RESUMO:** O glifosato é o herbicida de maior participação no mercado mundial, sendo utilizado intensivamente na agricultura há mais de 30 anos, principalmente devido ao custo relativamente baixo e à alta eficiência sobre diferentes espécies. Descreve-se um método para a síntese do glifosato na forma do sal isopropilamina enriquecido no isótopo estável do nitrogênio (<sup>15</sup>N). A síntese do herbicida-<sup>15</sup>N foi realizada utilizando-se da reação de fosfometilação com diaquil fosfito e glicina-<sup>15</sup>N. Os testes foram realizados em microescala e em quantidades equimolares. Foi possível obter um rendimento de 20%. Para testar a eficiência do glifosato sintetizado, realizou-se experimento em câmara de crescimento com a planta daninha *Lolium multiflorum* (azevém), a qual é eficazmente controlada segundo à bula dos produtos a base do sal de isopropilamina (glifosato). O delineamento experimental foi inteiramente casualizado, com quatro repetições, sendo os tratamentos provenientes de fontes de sal de isopropilamina na dose comercial recomendada do produto (720 g i.a. ha<sup>-1</sup>). Avaliou-se a produção de fitomassa fresca e seca, da parte aérea e radicular, aos 21 dias após a aplicação (DAA), e a toxicidade dos tratamentos aos 7, 14 e 21 DAA. As fontes de sal de isopropilamina empregados não diferiram, o que deverá permitir sua utilização como traçador isotópico.

**Palavras-chave:** *Lolium multiflorum*, sal de isopropilamina, isótopo estável

### Introduction

Glyphosate (C<sub>3</sub>H<sub>8</sub>NO<sub>5</sub>P), also known as N-phosphonomethyl glycine is the herbicide with the greatest participation in the world market. It is a non-selective, systemic, and post-emergent herbicide, included in the Brazilian toxicological classification (class IV - little toxic), as intensively used in weed control (Kertesz et al., 1995; Vencill, 2002). It is an easy-to-apply product, with relatively low cost and high effectiveness over different species (Giolo et al., 2005). Its herbicidal action

was first reported in 1971. It is sold in the form of aqueous concentrates of water-soluble salts, facilitating its application (Franz, 1985). Although there are more than 90 commercial brands of glyphosate worldwide (Heap, 1997), all of them present the same action mechanism, regardless of the used salts (Hartzler, 2006; Giolo et al., 2005).

When this molecule is used indiscriminately, the precautions associated with its potential environmental contamination must be carefully evaluated (Prata et al., 2000), since there has been evidence of the appearance

of resistant varieties, which will result in increased number of applications. Therefore, it becomes important to conduct studies dealing with the physicochemical properties of glyphosate, its interaction with water and soil, as well as the development of analytical tools that will allow its detection and quantification in natural samples.

Considering that the technology for the separation and production of glyphosate compounds enriched with the <sup>15</sup>N isotope is restrict, efforts were directed toward allowing the synthesis of the herbicide glyphosate, making an important tool available for evaluations of the fate of nitrogen in the soil-plant system, particularly in experiments conducted in the field, since no precautions with radioprotection would be needed, including the advantage of no generation of radiolabeled wastes.

In addition to synthesizing the compound in the laboratory, there is also the interested in evaluating whether the produced molecule has the same herbicidal potential as the commercial products. As a result, the effectiveness of sources of the isopropylamine salt was evaluated through the control of ryegrass (*Lolium multiflorum*), an annual, winter grass weed species that reproduces via seeds, characterized as typically sensitive to glyphosate. This species is frequently considered an undesired plant in wheat crops in Southern Brazil. (Roman et al., 2004).

## Material and Methods

### Production Methods

All steps required to produce <sup>15</sup>N-enriched glyphosate are represented schematically in Figure 1. A series of nitrogen compounds have already been obtained in recent years (Tavares et al., 2006; Bendassolli et al., 2002; Máximo et al., 2005). Early studies were conducted using reagents with natural isotopic abundance (0.366% <sup>15</sup>N atoms), considering that physical and chemical parameters would remain unchanged. Potential isotopic fractionation was evaluated from the production of glycine enriched with 0.9 or 6.7% <sup>15</sup>N atoms.

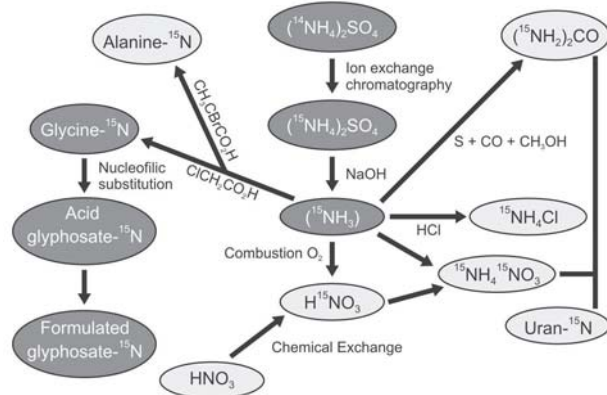


Figure 1 - Flow chart for the <sup>15</sup>N-enriched compounds production (highlighting the production steps related to glyphosate synthesis).

### Glyphosate synthesis from dialkyl phosphite and glycine (Patent No. 4,237,065 - 1980)

The mentioned patent reports to a process to obtain glyphosate in several steps, proposed by Ehrat and Dinhard (1980), aimed to use low-cost reagents that can be easily obtained in the market, as well as being exempt of purification methods in the final product. The reaction can be seen in the flow chart presented in Figure 2. According to the proposed synthesis, paraformaldehyde is dissolved in an alcohol under simultaneous depolarization, in the presence of a tertiary base (catalyst agent). Glycine is added to the obtained solution, forming hydroxymethylglycine as an intermediate product. This intermediate compound reacts immediately, under reflux, with added dialkyl phosphite. The resulting ester is saponified by the addition of a 30% NaOH solution (w/v); after this step, the tertiary base and the alcohol are removed. Finally, the obtained alkaline solution is acidified with the precipitation of N-phosphonomethyl glycine.

Under particular conditions, the amino acid glycine can be considered a simple amine soluble in alcohol or water. Glycine reacts extraordinarily as an amine in the Mannich reaction, involving formaldehyde, amine, and diethyl phosphite, resulting in phosphonomethyl amine. Nevertheless, with the conditions established in the patented invention the Mannich reaction does not produce an amine, but n-phosphonomethyl glycine. In this reaction, 1.5 mol paraformaldehyde are initially dissolved in 500 mL methanol, under simultaneous depolarization, in the presence of 1 mol triethylamine (tertiary base). One mol of glycine is added to the obtained solution, reacting with the formaldehyde from depolarization process, in a condition known as Mannich reaction (Holy et al., 1979). The solution is maintained at 70°C for 60 min, under reflux. Next, the solution is completed with 1 mol dialkyl phosphite by slowly dripping for 30 minutes. After agitation for 90 min, the initial solution gen-

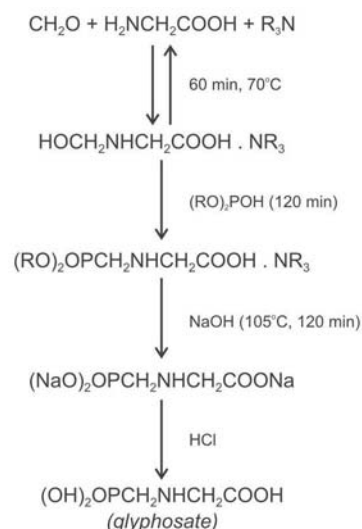


Figure 2 - Synthesis process to obtain acid glyphosate from dialkyl phosphite (Ehrat and Dinhard, 1980).

erates an ester, which is then saponified with 3.5 mol of a 30% sodium hydroxide solution (w/v). In this step, triethylamine and methanol are recovered by fractionated distillation at 70°C. Immediately after recovering the solvents, the system goes back into reflux at 105°C for a period of 2 h. Next, the obtained alkaline solution is acidified with concentrated hydrochloric acid until pH 1.5. The mixture is left to rest at 15°C for 3 to 4 h, during which glyphosate becomes crystallized and is then separated by filtration. A reaction yield of about 48 to 52% is achieved, producing a weight ranging from 81 to 88 g.

The preliminary tests were conducted with three replicates, at a reduced scale of 1:100, using glycine with natural isotopic abundance (0.366%  $^{15}\text{N}$  atoms). After establishing the best experimental conditions, two tests were made, starting from glycine with isotopic abundances of 6.70 and 0.90% in  $^{15}\text{N}$  atoms, in order to evaluate potential isotopic fractionation during the process. The system used in the synthesis process was identical to the one illustrated in Figure 2. The isotopic determination in the synthesized glyphosate was made using an ANCA-SL 20/20 mass spectrometer.

After synthesis, the determinations of interest were made in order to confirm the presence of the herbicide, based on fusion tests, infrared spectroscopy, and total nitrogen content determination by mass spectrometry. The synthesized samples were compared against the Sigma standard (p.a. grade). Another possibility to confirm the presence of the herbicide would be to start by nuclear magnetic resonance in the solid phase (Li and Evans, 1995).

### Production of formulated Glyphosate

One of the steps in obtaining the formulated herbicide consists in the synthesis of the glyphosate salt, which can be achieved by neutralizing the acid glyphosate with a base, such as monoisopropylamine. This step of the production process is referred to as pre-mixture or first formulation step, containing a 62% acid concentration.

Acid (or solid) glyphosate is relatively insoluble in water, making its application in the field not viable. Because of this, glyphosate is typically formulated in the form of salts, making the molecule soluble. This particular trait allowed several salts to be obtained with a commercial objective in mind. Studies comparing the efficiency of various salts have demonstrated that isopropylamine has an increased phytotoxicity in plant. However, the salt, in association with a surfactant, increases absorption by the plant (Duke, 1988).

In this investigation, it was chosen to obtain the herbicide under the same conditions of the commercial product. To achieve that, the acid glyphosate obtained in the previous step was submitted to amination with 13.5% oleamide monoisopropylamine (MIPA), in association with 13% tensioactive agent (surfactant – ultramine 200).

Ultramine 200 is an ethoxylated fatty amine with moisturizing characteristics. It is recommended for agrochemical formulations in its pure or combined form as a moisturizing or solubilizing agent in soluble concentrates of water-based herbicidal formulations, with a spreading effect and as an aid to improve penetration.

Samples obtained from this chemical reaction were analyzed by HPLC with a fluorescence detector, using middle-infrared spectroscopy for solid and liquid samples. Two assays were conducted to determine the  $^{15}\text{N}$  isotopic dilution in the formulated glyphosate, using glycine enriched at 6.7 or 0.9% in  $^{15}\text{N}$  atoms as precursors.

The contribution of the nitrogen source from the acid glyphosate and from amination (MIPA), as well as from the tensioactive agent (surfactant) can be determined in the final product based on the isotopic mass balance. Figure 3 illustrates schematically how the glyphosate formulation is obtained from the MIPA acid added of Surfactant (S). The isotopic mass balance equation can be represented by Equation 1.

$$A f_{(AG)}^N + B f_{(M+S)}^N = C (f_{(AG)}^N + f_{(M+S)}^N) \quad (1)$$

where: A, B, and C: isotopic abundances (% in  $^{15}\text{N}$  atoms) in the acid glyphosate, MIPA + Surfactant mixture, and formulated glyphosate, respectively;  $f_{(AG)}^N$  and  $f_{(M+S)}^N$  = nitrogen fraction in the final product, from the acid glyphosate and the MIPA and Surfactant mixture, respectively.

Considering that the MIPA (oleamide monoisopropylamide) and Surfactant (ultramine) mixture have natural isotopic abundance (0.366% in  $^{15}\text{N}$  atoms), and also the nitrogen mass balance given by equation 2, equation 3 can be obtained.

$$f_{(AG)}^N + f_{(M+S)}^N = 100 \text{ or } f_{(M+S)}^N = 100 - f_{(AG)}^N \quad (2)$$

$$f_{(AG)}^N = [(C - 0.366)/(A - 0.366)] \cdot 100 \quad (3)$$

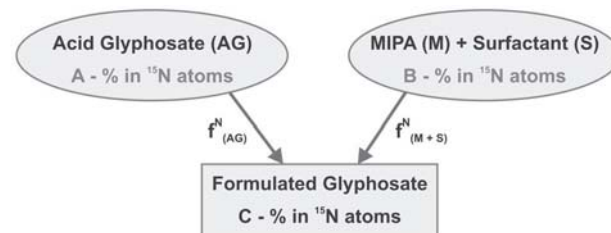


Figure 3 - Flow chart showing how the formulated glyphosate is obtained (acid glyphosate + MIPA + Surfactant).  $f_{(AG)}^N$  = N fraction from the acid glyphosate in the final product;  $f_{(M+S)}^N$  = N fraction from the MIPA and Surfactant mixture in the final product; A, B, and C correspond, respectively, to  $^{15}\text{N}$  isotopic abundance (% in  $^{15}\text{N}$  atoms) in the acid glyphosate, MIPA + Surfactant, and formulated glyphosate (final product).

### Effectiveness test for the synthesized product

The experiment was performed in a growth chamber at 24/19°C day/night with 250  $\mu\text{mol m}^{-2}$  s photosynthetically active radiation, and a 16-h photoperiod. The relative humidity oscillated between 60 and 80%. These conditions approximately correspond to the normal ones found in the field when glyphosate is applied. The weed *Lolium multiflorum*, which is susceptible to the herbicide glyphosate was used in this study. The seeds were initially placed to germinate in trays. When 50% + 1 plants reached the two-leaf phenological stage, they were transplanted to the experimental plots, at a density of three plants per pot. The plots consisted of 0.5 L-pots filled with commercial substrate and fertilized with a dose proportional to the application of 500 kg ha<sup>-1</sup> of the commercial formulation 10-10-10 (N-P<sub>2</sub>O<sub>5</sub>-K<sub>2</sub>O), daily irrigated. The experimental design was completely randomized with four replicates.

The treatments consisted of isopropylamine salt sources at the commercial dose recommended for the control of *Lolium multiflorum*, corresponding to 720 g a.i. ha<sup>-1</sup> plus a control treatment. The treatments were applied when the plants had four completely-expanded leaves, on average. The sprays were made in a closed application chamber, using a flat-fan nozzle, with the jet calibrated at a 50 cm height from the target surface, with a relative mix volume in the order of 200 L ha<sup>-1</sup>. After the herbicide had been applied, the pots were returned to the growth chamber and they were irrigated on the following day, ensuring the leaf absorption of the herbicide.

Treatment toxicity evaluations were made at 7, 14, and 21 days after application (DAA), using a percentage scale, in which a zero rating meant no damage effect to the plant and a 100 rating represented complete plant suppression or death. After the last evaluation, the above-ground part and root of the plants were collected to determine green and dry mass yield. All sampled materials were dried in a forced air circulation oven at 60°C for 72 h, allowing dry phytomass determination.

The F test on analysis of variance was applied to the obtained data. The means for green and dry phytomass data and weed control treatment toxicity values were compared by the Tukey test ( $p < 0.05$ ).

## Results and Discussion

### Glyphosate synthesis from dialkyl phosphite and glycine (Patent No. 4,237,065 - 1980)

The evaluation of the parameters involved in the production of glyphosate from synthesized glycine and dialkyl phosphite, a process established by Erhat & Dinhard, (1980), showed that this is the most promising alternative for glyphosate synthesis under the operational conditions available at the Stable Isotopes Laboratory of CENA.

To evaluate the N percentage in the herbicide, synthesis assays were performed using glycine p.a. (Sigma) and the laboratory-synthesized glycine as the starting

points. The results for these tests are presented in Table 1. The glyphosate synthesized from dialkyl phosphite had N content (%) and isotopic abundance (% in <sup>15</sup>N atoms) similar to the Sigma and the standards (commercial standard). The natural isotopic abundance of the heavy nitrogen isotope is  $0.366 \pm 0.004$ . According to the infrared analyses, the results were similar between synthesized samples and the Sigma standard.

Under the tested conditions, a yield about 26% was obtained as compared to the 50% yield predicted in the literature (Table 2). The use of synthesized or natural glycine did not influence the final result.

### Production of formulated Glyphosate

In relation to the samples analyzed by HPLC via a FMOC derivative, a similar retention time between the standard (27.60 min), the synthesized sample with natural abundance (27.56 min), and the synthesized <sup>15</sup>N-labeled sample (27.55) was observed. When compared against the standard, the results apparently indicate the same substance. The results can be seen in Figures 4, 5, and 6, respectively, for the standard, the synthesized sample with natural abundance, and the synthesized <sup>15</sup>N-labeled sample.

At the same time, analyses were made using middle infrared with raw samples. When the glyphosate sample with natural abundance, the <sup>15</sup>N-labeled sample, and the formulated standard provided (commercial product) were compared, the emitted peaks practically over-

Table 1 - N percentage, m/m, and isotopic abundance, % in <sup>15</sup>N atoms, in glyphosate samples.

Glyphosate	Nitrogen %	Isotopic abundance % in <sup>15</sup> N atoms
T1	8.11 ± 0.01	0.363 ± 0.002
T2	8.24 ± 0.13	0.363 ± 0.002
Sigma standard	8.50 ± 0.21	0.366 ± 0.001
Standard acid (commercial)	8.60 ± 0.51	0.365 ± 0.003

T1 = glyphosate obtained from glycine p.a.; T2 = glyphosate obtained from synthesized glycine in Stable Isotopes Laboratory.

Table 2 - Mass and reaction yield for glyphosate synthesized from glycine and dialkyl phosphate.

Test	Mass g	Yield %
T1	0.4194	25.73
T2	0.4314	26.46
T3	0.4439	27.23
T4	0.4098	25.14
Mean ± s.d.	0.426 ± 0.01	26.14 ± 0.70

T1 = acid glyphosate obtained from glycine synthesized with 6.7% in <sup>15</sup>N atoms; T2, T3, and T4 = acid glyphosate obtained from glycine synthesized in this study with natural isotopic abundance (0.366% in <sup>15</sup>N atoms).

Table 3 - Isotopic and N fraction data in the final product (formulated glyphosate).

Precursor <sup>(1)</sup>	Compound	Isotopic abundance (% in <sup>15</sup> N atoms)	N contribution in t <sup>bc</sup> product (%) <sup>(4)</sup>	
			f <sup>N</sup> (AG)	f <sup>N</sup> (M+S)
Glycine (6.70% in <sup>15</sup> N atoms)	Acid glyphosate	6.72 <sup>(2)</sup>	36.6	63.4
	Formulated glyphosate	2.64 <sup>(3)</sup>		
Glycine (0.90% in <sup>15</sup> N atoms)	Acid glyphosate	0.87	34.5	65.5
	Formulated glyphosate	0.54		

<sup>(1)</sup>previously enriched glycine used for acid glyphosate synthesis; <sup>(2)</sup>and <sup>(3)</sup>values for A and C in equation (4); <sup>(4)</sup>obtained from equation (4), corresponding to f<sup>N</sup>(AG) and f<sup>N</sup>(M+S).

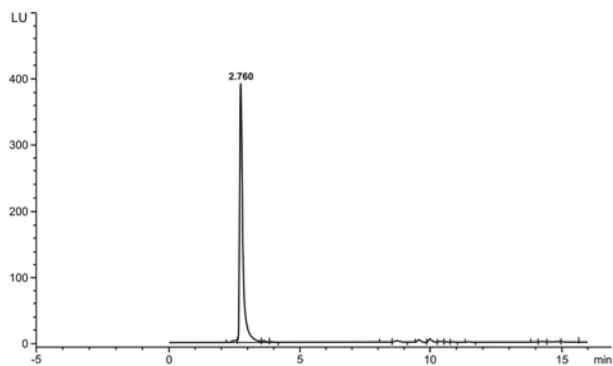
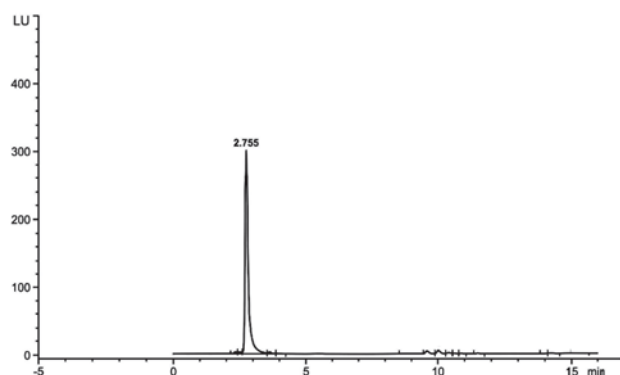
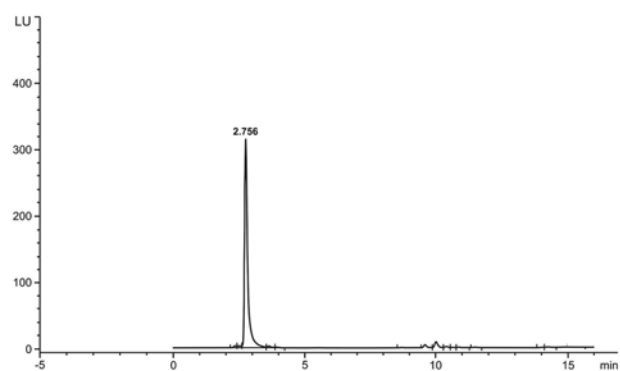
Figure 4 - Formulated glyphosate standard 480 g L<sup>-1</sup> (commercial standard).

Figure 5 - Formulated synthesized glyphosate sample with natural isotopic abundance.

Figure 6 - <sup>15</sup>N-labeled synthesized glyphosate sample.

lapped. These tests were conducted using a Thermo Nicolet IR200 spectrometer with Fourier transform (FTIR).

Regarding the <sup>15</sup>N-enriched glyphosate samples, in which glycine previously enriched in the isotope was used (6.70 or 0.90% in <sup>15</sup>N atoms), it was possible to obtain acid and formulated glyphosate with the isotopic characteristics listed in Table 3. The same Table presents the isotopic values and contributions of both nitrogen sources (N-acid glyphosate and N-MIPA + Surfactant) in the formulated product (acid glyphosate + MIPA + Surfactant). The contribution from each nitrogen source was obtained from isotopic balance equations.

There was no isotopic fractionation in the conversion from <sup>15</sup>N-glycine into acid glyphosate, and the isotopic dilution while obtaining the formulated glyphosate was practically the same of the assays that used acid glyphosate containing 6.70 or 0.90% in <sup>15</sup>N atoms (Table 3). The <sup>15</sup>N isotopic dilution in the final (formulated) product due to the MIPA and Surfactant introduction does not represent a problem in the use of the herbicide as tracer, considering that what must be taken into account is the active principle enrichment.

### Effectiveness Test

Equal toxicity response injuries were obtained among the tested treatments, only differing from the control, in the evaluated periods (Figure 7). Consequently, in relation to the control effectiveness parameter against the weed *Lolium multiflorum* by different isopropylamine salt sources, all of them were not statistically different. The variation coefficient found in the variable above-ground for fresh and dry mass were 22.47% and 24.28%, respectively; in the root for fresh and dry mass were 46.74% and 25.68%, respectively, and in the control percentage was 33.86%.

For each treatment, a general decrease in fresh of the root system was observed relative to its corresponding control. These results agree with those obtained in the toxicity evaluation (Figures 7 and 8). This pattern was also maintained for dry mass of the above-ground part of *Lolium multiflorum* (Figure 8). The fresh and dry phytomass for the root system had the same pattern observed for the above-ground part of this plant. However, smaller decreases than those for the above-ground part of *Lolium multiflorum* were observed relative to the control. (Figure 8).

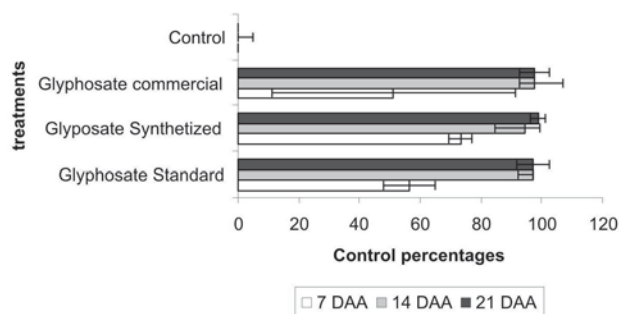


Figure 7 - Control percentages of the weed *Lolium multiflorum* and standard deviations, presented by different isopropylamine salts at 7, 14, and 21 days after application (DAA).

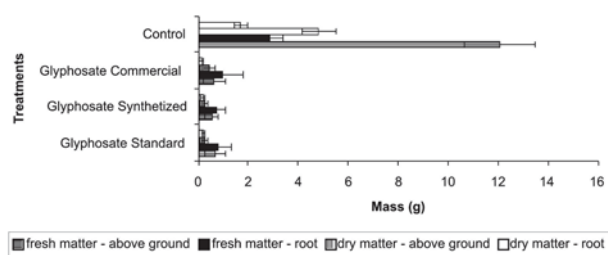


Figure 8 - Mean values and standard deviations for green and dry mass of the above-ground part and root system of the weed *Lolium multiflorum*, presented by different isopropylamine salts at 21 days after application (DAA).

## Conclusions

The method that best suited the interests of the proposed study was the one that employed dialkyl phosphite, allowing acid glyphosate to be obtained with a yield of about 25%, corresponding to 0.44 g glyphosate in the acid form, starting from 0.75 g <sup>15</sup>N-glycine. Regarding the glyphosate formulated in the isopropylamine salt form, an isotopic dilution of about 60% was obtained. However, in the production of highly enriched glyphosate (80 – 90% in <sup>15</sup>N atoms) this problem can be overcome. As a result, active principle-labeling becomes the factor that should be taken into consideration. The herbicide can be obtained under the same conditions as the commercial product, including a lack of formation of undesired impurities in the synthesis pathway.

## Acknowledgements

The authors thank FAPESP for the financial support.

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Received September 10, 2007

Accepted August 21, 2009