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SOLVENT BASED SWITCHING OF THE PHOTOPHYSICAL PROPERTIES OF TRANSITION METAL COMPLEXES

Wassim Zuhair Alsindi, M.Sci.

Thesis submitted to the University of Nottingham for the degree of Doctor of Philosophy
April 2007
Declaration

Except where specific reference is made to other sources or collaborators, the work presented in this Thesis is the original work of the author. It has not been submitted, in whole or in part, for any other Degree.

Signed: Wassim Alsindi

Date: 01/06/07
The work presented in this Thesis describes the modular design and spectroscopic study of polynuclear systems based on ruthenium (II) and rhenium (I) complexes. A combination of UV/vis, luminescence and TRIR spectroscopies, electrochemistry, spectroelectrochemistry and conformational analysis have been employed to understand the electronic structure of the ground and excited states of these compounds.

Chapter 1 gives an introductory background to this Thesis. An overview of transition metal photophysics and excited states, and the typical spectroscopic and electrochemical techniques used in their study is presented. Previous studies of the ground and excited state properties of the complexes \([\text{Ru(bpy)}_3]^2^+\) and \([\text{ReCl(CO)}_3(bpy)]\) which are used as supramolecular building blocks in this Thesis are presented and a number of relevant studies of supramolecular systems are described.

Chapter 2 contains a study extending the known family of \([\text{Ru(CN)}_6(NN)]^3^-\) complexes and describes their unique advantages over \([\text{Ru(bpy)}_2]^3^-\). The results obtained are discussed alongside previous studies. This completes the introduction of the molecular building blocks used in Chapters 3 and 4.

Chapter 3 details a study of through-space PEnT in bimetallic systems constructed from the complexes introduced in Chapters 1 and 2, bridged by a saturated alkyl linker between bpy ligands on either metal. This Chapter demonstrates the solvent-switchable nature of the direction and gradient of PEnT, using ps-TRIR spectroscopy to directly probe these processes in real time.

Chapter 4 describes a study of bimetallic systems bridged by conjugated the ligand 2,2'-bipyridine. Monometallic, homobimetallic and heterobimetallic systems are studied and questions arising from limitations of previous studies are addressed. In particular ps-TRIR spectroscopy gives new insight into the numerous ultrafast processes occurring.

Chapter 5 summarises the achievements of this Thesis and suggests promising directions for extending this work in the future.

Chapter 6 describes the experimental and theoretical techniques used in this Thesis.
Acknowledgements

Firstly I would like to extend sincere thanks and appreciation to my supervisor Professor Mike George for his guidance, support and encouragement over the course of my studies. I have also received a great deal of help from past members of the Nottingham Photochemistry and Spectroscopy group in the early stages of this work, notably Dr. Peter Portius, Dr. Jason Hyde, Dr. Julia Weinstein, Dr. Jixin Yang and Dr. Marina Kuimova. Thanks also go to Dr. Stephen Davies for his help with electrochemistry and Dr. Jon McMaster for his assistance with computational chemistry.

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This project had a strong collaborative focus and I have worked in tandem with Tim Easun and Professor Mike Ward at the University of Sheffield who have been excellent co-workers. The Ward group’s expertise in synthetic chemistry has complemented our spectroscopic and photophysical skills and this project would not have been possible had it not been for their preparation of all the bespoke complexes in this Thesis. Thanks go to Tim Easun for being such a motivated collaborator and Nina Depperman and Dr. Juan-Manuel Herrera for their role in the preparation of some of the complexes.

This work would have not been possible without access to the PIRATE spectrometer at the Lasers for Science Facility in the Central Laser Facility at the Rutherford Appleton Laboratory. I would like to thank Tony Parker, Pavel Matousek, Emma Belcher and particularly Kate Ronayne and Mike Towrie for the amount of time and effort put into keeping our experiments running late into the night and helping us to unravel the often mysterious spectra that we obtained. Dr. Xue-Zhong Sun and Tim Easun also deserve
praise for their commitment and dedication to getting the most out of our limited experimental time, often working in the laboratory until past bedtime!

I would also like to thank Dr. June McCombie, Dr. Helen Fraser and Professor Ewine van Dishoeck for giving me the opportunity to undertake my Masters research project in the wonderful surroundings of de Sterrewacht (Astronomical Observatory), working in the Raymond and Beverly Sackler Laboratory for Experimental Astrophysics, University of Leiden. The semester I spent there in early 2003 was a fantastic prelude to my Ph.D. studies and provided the perfect opportunity to ‘cut my teeth’ at a world leading institution.

A special thank you goes to Katharine T. Adcock for her support and understanding throughout the writing process and praise also goes to my business partner Darren Gower for his patience over the last six months, as I haven’t been pulling my weight with the day-to-day workload of running our company!

Finally I would like to thank my family and in particular my parents, Zuhair and Sahar Alsindi who have made so many sacrifices for their children’s future. My thoughts are perpetually with my extended family, who have suffered war, displacement and hardship since I began this work in 2003. The world may have forgotten about you and the millions of Iraqi diaspora but I have not.

**List of Abbreviations**

<table>
<thead>
<tr>
<th>Symbol</th>
<th>Definition</th>
</tr>
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<tbody>
<tr>
<td>ΔE</td>
<td>Energy difference</td>
</tr>
<tr>
<td>ΔOD</td>
<td>Change in absorbance</td>
</tr>
<tr>
<td>λ</td>
<td>Wavelength</td>
</tr>
<tr>
<td>λ_{abs}</td>
<td>Absorption wavelength</td>
</tr>
<tr>
<td>λ_{em}</td>
<td>Emission wavelength</td>
</tr>
<tr>
<td>λ_{ex}</td>
<td>Excitation wavelength</td>
</tr>
<tr>
<td>λ_{max}</td>
<td>Wavelength maximum of peak</td>
</tr>
<tr>
<td>μs</td>
<td>Microsecond (10^{-6} s)</td>
</tr>
<tr>
<td>ν</td>
<td>Wavenumber</td>
</tr>
<tr>
<td>τ</td>
<td>Excited state lifetime</td>
</tr>
<tr>
<td>φ</td>
<td>Luminescence quantum yield</td>
</tr>
<tr>
<td>χ^{2}</td>
<td>Goodness of fit test</td>
</tr>
<tr>
<td>χ^{1}</td>
<td>Singlet (excited state)</td>
</tr>
<tr>
<td>χ^{3}</td>
<td>Triplet (excited state)</td>
</tr>
<tr>
<td>(NN)</td>
<td>α-Diimine</td>
</tr>
<tr>
<td>§</td>
<td>Section/Chapter</td>
</tr>
<tr>
<td>Å</td>
<td>Angström (10^{-10} m)</td>
</tr>
<tr>
<td>A / Abs</td>
<td>Absorbance</td>
</tr>
<tr>
<td>A. N.</td>
<td>Gutmann Acceptor Number</td>
</tr>
<tr>
<td>BL</td>
<td>Bridging Ligand</td>
</tr>
<tr>
<td>CV</td>
<td>Cyclic Voltammetry</td>
</tr>
<tr>
<td>DCM</td>
<td>Dichloromethane</td>
</tr>
<tr>
<td>DFT</td>
<td>Density Functional Theory</td>
</tr>
<tr>
<td>DMF</td>
<td>Dimethyl formamide</td>
</tr>
<tr>
<td>DMSO</td>
<td>Dimethyl sulfoxide</td>
</tr>
<tr>
<td>D. N.</td>
<td>Gutmann Donor Number</td>
</tr>
<tr>
<td>Et</td>
<td>Ethyl</td>
</tr>
<tr>
<td>FTIR</td>
<td>Fourier Transform Infrared</td>
</tr>
<tr>
<td>H-bonding</td>
<td>Hydrogen-bonding</td>
</tr>
<tr>
<td>HAT</td>
<td>Hexaaza-triphenylene</td>
</tr>
<tr>
<td>HOMO</td>
<td>Highest Occupied Molecular Orbital</td>
</tr>
<tr>
<td>IC</td>
<td>Internal conversion</td>
</tr>
<tr>
<td>IL</td>
<td>Intra-ligand</td>
</tr>
<tr>
<td>IR</td>
<td>Infrared</td>
</tr>
<tr>
<td>IRF</td>
<td>Instrument Response Function/Time</td>
</tr>
<tr>
<td>ISC</td>
<td>Intersystem Crossing</td>
</tr>
<tr>
<td>IVCT</td>
<td>Inter-Valence Charge Transfer</td>
</tr>
<tr>
<td>LC</td>
<td>Ligand-centred</td>
</tr>
<tr>
<td>LMCT</td>
<td>Ligand-to-Metal Charge Transfer</td>
</tr>
<tr>
<td>LUM</td>
<td>Luminescence</td>
</tr>
<tr>
<td>LUMO</td>
<td>Lowest Unoccupied Molecular Orbital</td>
</tr>
<tr>
<td>Acronym</td>
<td>Description</td>
</tr>
<tr>
<td>---------</td>
<td>-------------</td>
</tr>
<tr>
<td>M</td>
<td>Metal centre</td>
</tr>
<tr>
<td>MC</td>
<td>Metal-Centred</td>
</tr>
<tr>
<td>MCT</td>
<td>Mercury Cadmium Telluride</td>
</tr>
<tr>
<td>Me</td>
<td>Methyl</td>
</tr>
<tr>
<td>MLCT</td>
<td>Metal-to-Ligand Charge Transfer</td>
</tr>
<tr>
<td>NMR</td>
<td>Nuclear Magnetic Resonance</td>
</tr>
<tr>
<td>ns</td>
<td>Nanosecond ($10^{-9}$ s)</td>
</tr>
<tr>
<td>OTE</td>
<td>Optically Transparent Electrode</td>
</tr>
<tr>
<td>OTTLE</td>
<td>Optically Transparent Thin Layer Electrode</td>
</tr>
<tr>
<td>PEnT</td>
<td>Photoinduced Energy Transfer</td>
</tr>
<tr>
<td>PET</td>
<td>Photoinduced Electron Transfer</td>
</tr>
<tr>
<td>ps</td>
<td>Picosecond ($10^{-12}$ s)</td>
</tr>
<tr>
<td>PIRATE</td>
<td>Picosecond Infrared Absorption Transient Excitation</td>
</tr>
<tr>
<td>PMT</td>
<td>Photo-Multiplier Tube</td>
</tr>
<tr>
<td>PPN$^+$</td>
<td>Bis(triphenylphosphine)iminium</td>
</tr>
<tr>
<td>r</td>
<td>Förster interchromophoric distance</td>
</tr>
<tr>
<td>RAL</td>
<td>Rutherford Appleton Laboratory</td>
</tr>
<tr>
<td>SSDA</td>
<td>Second Sphere Donor-Acceptor</td>
</tr>
<tr>
<td>T2D-IR</td>
<td>Transient Two-Dimensional Infrared</td>
</tr>
<tr>
<td>TCSPC</td>
<td>Time-Correlated Single Photon Counting</td>
</tr>
<tr>
<td>TRIR</td>
<td>Time Resolved Infrared</td>
</tr>
<tr>
<td>UV</td>
<td>Ultraviolet</td>
</tr>
<tr>
<td>UV/vis</td>
<td>Ultraviolet-visible</td>
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1.1 Supramolecular Photoscience

1.1.1 Overview and Excited States

Supramolecular chemistry is ‘chemistry beyond the molecule’.\(^1\) A supramolecular system consists of molecular components, just as molecules consist of atomic components. Supramolecular systems may possess properties that are not directly related or derived from their constituent units due to the intermolecular interactions taking place in the supramolecular structure. This can have significant effects on their photoreactivity as photoinduced processes between the molecular components become possible.

It is important to distinguish between photophysical and photochemical processes as these words are often erroneously used interchangeably. Photophysics refers to a process involving the absorption, emission or radiationless quantum transition / event in which there is no chemical change taking place: in other words the process involves a change in electronic structure (the formation of an excited state) without a change in chemical formula of the analyte. Conversely photochemical processes involve a change in chemical formula, which may be reversible or irreversible depending on the experimental conditions used. In this case the change in electronic structure induced by absorption, emission or quantum transition of a photon is accompanied by a chemical reaction. This Thesis is solely concerned with the excited state photophysics of the complexes prepared and studied in Chapters 2-4 and it should be emphasised that no (desired) chemical reactions are taking place.

A Jabłoński diagram may be drawn to give an overview of the events in a system from the formation of an excited state to its deactivation and return to the ground state electronic configuration. Figure 1.1.1.1 contains a generalised form of this diagram which is relevant for the majority of isolated metal centres in this Thesis.
With reference to Figure 1.1.1.1, the formation, evolution and subsequent decay of excited state species may be described as follows. Absorption of a photon $h\nu_A$ by a ground state system $S_0$ leads to formation of a singlet excited state $S_1$ or $S_2$, with excitation into a "hot" excited vibrational energy level ($v > 0$) also being commonplace. Radiationless relaxation of the electronically excited state to the lowest energy singlet species (usually $S_1$) takes place rapidly. This internal conversion typically occurs on a sub-picosecond timescale. Following this process the relaxation of vibrationally excited states to the $v = 0$ energy level (vibrational cooling) may then take place, typically on a timescale of a few picoseconds. This $S_1$ excited state may be metastable (particularly in organic systems) and its deactivation may either be radiative (via fluorescence or more generally defined as luminescence from a singlet state) or non-radiative.

In systems containing transition metal atoms, the $S_1$ state may rapidly convert into a $T_1$ (triplet) excited state (vide infra) via a process known as intersystem crossing. As with $S_1$ states, a $T_1$ excited state may return to the $S_0$ ground state through radiative (luminescence from a triplet state) or non-radiative deactivation. The words 'fluorescence' and 'phosphorescence' should be used with care when...
referring to transition metal complexes, as mixing of singlet and triplet excited states may occur and the effect of spin-orbit coupling may lead to relaxation of selection rules (§1.1.2). Spin-orbit coupling may have such an effect that pure singlet or triplet emission may not necessarily occur, and the emission is best referred to as luminescence.

1.1.2 Monometallic Complexes

The majority of photophysically interesting coordination compounds in the present day literature contain a metal centre (usually ReI, RuII, RhI, IrI or PtII) and a N-heterocyclic aromatic ligand, typically a polypyridyl such as 2,2'-bipyridine (bpy). The archetypal complex in this particular paradigm is [Ru(bpy)3]2+ which is discussed in §1.4.1.1. In photophysical investigations of such complexes, excitation wavelengths are chosen to coincide with the absorption energy of a charge transfer transition, usually metal-to-ligand charge transfer (MLCT). This process involves the transition of the excited electron from the metal based HOMO orbital (normally a dπ orbital) to the ligand-based LUMO orbital (usually a delocalised pπ* antibonding orbital) which is typically located on a polypyridyl ligand.

The formation of a 3MLCT (triplet MLCT) excited state leads to large changes in the charge distribution of electron density in the excited state of the complex. In general 3MLCT excited states are relatively long lived, with lifetimes in the range of nanoseconds to microseconds, facilitating their study with time-resolved spectroscopic methods on the fast and ultrafast timescales. This is partially due to their triplet nature, i.e. the initially formed singlet excited state undergoes a ‘spin flip’ to generate the energetically favoured triplet species with both unpaired electrons possessing the same spin orientation. This corresponds to the intersystem crossing process in Figure 1.1.1.1. Intersystem crossing is facilitated in photophysically active metal complexes due to efficient spin-orbit coupling mechanisms relaxing spin conservation selection rules owing to the presence of heavy metal atoms (vide supra). The lifetime of 3MLCT excited states is also usually longer than IL (intraligand, transitions within the ligand manifold) or MC
(metal centred) excited states as a result of the 'spin-forbidden' transition required to return a triplet excited state to its singlet ground state electronic configuration.

1.1.3 Bimetallic and Polynuclear Complexes

In complexes containing more than one metal centre an intermetallic electronic interaction can occur after one metal centre is initially promoted to an excited state by the absorption of a photon. These processes are photoinduced interactions in chromophore-quencher systems. The excited state of one component can undergo photoinduced electron-transfer (PET) to or from the adjacent metal centre, resulting in charge-separated states ('P' and Q' or 'P' and Q') in Figure 1.1.3.1); or photoinduced energy-transfer (PEt) can occur such that the secondary metal fragment ends up in an excited state (Q') and the first one returns to its ground state (P, Figure 1.1.3.1).

Figure 1.1.3.1: A supermolecule consisting of a photosensitiser (P) and a quencher (Q) may undergo excitation following photon absorption (P'). PET leads to charge-separation (only one possible product shown) and PEt leads to the formation of an excited state on the quencher (Q') component.

The absorption of a photon leads to the formation of a localised excited state. Photoinduced Electron Transfer (PET) from one metal centre to another can take place via a radiationless process, providing that the potential energy surfaces of the excited state system (e.g. P'-Q in Figure 1.1.3.1) and the charge-separated system
(P'-Q') overlap sufficiently and the charge-separated state is thermodynamically more stable than the excited state.\textsuperscript{5}

Photoinduced Energy Transfer can occur in two ways: the Förster-type mechanism\textsuperscript{6} and the Dexter-type mechanism.\textsuperscript{7} The Förster mechanism operates at long-ranges and is based on Coulombic interactions. Energy is exchanged through space via coupling of donor and acceptor transition dipoles and is efficient when the emission spectrum of the donor is coincident with the absorption spectrum of the donor and there is good orientational alignment of the two transition dipoles. The rate and efficiency of Förster-type PEnT exhibits a sensitive dependence on the distance between the donor and acceptor chromophores:\textsuperscript{8}

\[
k_T = \frac{(R_0 / r)^6}{\tau_D} = A \frac{\kappa^2 k_D}{r^6} J(E)
\]

where \(k_T\) is the rate of energy transfer, \(\tau_D\) is the decay time of the donor in the absence of the acceptor, \(R_0\) is the Förster distance (the distance at which PEnT efficiency is 50%), \(r\) is the interchromophoric distance, \(\kappa^2\) is the dipole-dipole orientational factor, \(k_D\) is the rate of excited state decay of the donor species in the absence of the acceptor and \(J\) is the overlap integral term.

The Dexter mechanism is based on exchange interactions and operates at shorter distances. Direct orbital overlap is necessary for efficient energy transfer via the Dexter mechanism. The prevalent mechanism of energy transfer can be influenced by the design of the supramolecular system.\textsuperscript{5}

PEnT is of interest in the development of luminescent sensors and systems designed to harvest visible light. Making these interactions switchable, such that an interaction can be switched 'on' or 'off' reversibly and controllably under the influence of some external stimulus, is of great interest to provide mechanisms to control electron and energy transfer.\textsuperscript{9} This Thesis explores the use of solvatochromism and bridging ligand properties such as conjugation, structure and geometry to control photoinduced processes in polynuclear systems.
1.2 Solvatochromism

1.2.1 General Solvent-Solute Effects in UV/vis Absorption Spectra

Solvatochromism is defined as the change in intensity and/or position of a UV/visible absorption profile following a change in the constitution of the solvent medium.\(^\text{10}\) It is important to remember that UV/vis absorption spectroscopy probes the Franck-Condon electronic excited states populated instantaneously (\(< 10^{-14} \text{ s}\)) following absorption of a photon. It does not give any information on molecular dynamics and is only influenced by the immediate solvent sphere during the photon absorption event.

Two general trends are observed in solvent effects on absorption transition energies: positive solvatochromism and negative solvatochromism (Figure 1.2.1.1). Positive solvatochromism is a decrease in transition energy (or shift to longer wavelength) as the solvent ‘polarity’ increases (see §1.2.3). This occurs because the electronic transition accompanying the absorption of a photon causes a redistribution of electron density in the excited state, leading to an increase in the dipole moment. In coordination compounds this is commonly observed for MLCT and (to a lesser extent) intraligand (IL) \(\pi-\pi^*\) transitions. The greater ability of the surrounding solvent shell to stabilise the increased polarity of the excited state, relative to the ground state, leads to red-shifting of the absorption energy.

![Solvatochromism Diagram](image)

**Figure 1.2.1.1:** The effects of solvent-solute interactions on transition energies in optical absorption spectroscopy; a) is positive and b) is negative solvatochromism. Adapted from Ref.\(^\text{10}\)
Negative solvatochromism is an increase in the transition energy (or shift to shorter wavelength) as the solvent ‘polarity’ increases. Here the electronic transition leads to a decrease in the dipole moment of the excited state. This is observed in systems with n-\(\pi^*\) transitions.

### 1.2.2 General Solvent-Solute Effects in Luminescence Spectra

Emission spectroscopy probes much longer timescales than absorption spectroscopy (see §1.3). The solvent molecules align their dipoles with the excited state’s dipole by rotating. This lowers the energy of the excited state before emission of a photon and relaxation to the ground state (Figure 1.2.2.1).\(^\text{10}\) The more ‘polar’ the solvent, the greater the excited state stabilisation. Following photon emission the solvent dipoles are still aligned to that of the excited state. Thus the energy of the state formed following photon emission has a higher energy than the initial ground state. These factors lead to emission energies being lower than absorption energies and emission spectra being red-shifted in polar solvents.

![Figure 1.2.2.1: Solvent effects on luminescence spectroscopy. The dipole moments of the molecule (large arrow) and the solvent shell (small arrows) are shown. a) Absorption of a photon, b) solvent relaxation and c) emission. In terms of polarity Solv 3 > 2 > 1. Adapted from Ref.\(^\text{10}\).](image-url)
1.2.3 Specific Solvent-Solute Effects

Significant electronic interactions between the solute and the solvent shell can take place in certain situations. For example sites available for H-bonding or coordination may be present on the ligands, allowing Second Sphere Donor-Acceptor (SSDA) interactions to take place between the solvent shell and ligands. These specific effects have been studied in ammine and cyanide complexes,\textsuperscript{11,12} and result in spectroscopic properties which depend on the properties of the solvent.

Several systems have been introduced to quantify the ability of a solvent to participate in donor-acceptor interactions, including the Donor Number (D. N.) and Acceptor Number (A. N.) introduced by Gutmann.\textsuperscript{13} The D. N. is based on the molar reaction enthalpy in 1,2-dichloroethane for the formation of a 1:1 adduct between SbCl$_5$ and the added solvent as donor. It gives a measure of the ability of solvent molecules to donate an electron pair to SbCl$_5$, and is used to describe H-bonding with (Lewis acidic) ammine complexes (Figure 1.2.3.1a).\textsuperscript{14}

The Acceptor Number is taken as the chemical shift of the $^{31}$P resonance of Et$_3$PO in the solvent relative to two reference measurements. Et$_3$PO in hexane is assigned A. N. = 0 and the Et$_3$PO:SbCl$_5$ adduct is taken as A. N. = 100. The Acceptor Number measures the ability of solvent molecules to act as electron-pair acceptors and is used to quantify specific effects with cyanide complexes (Figure 1.2.3.1b).\textsuperscript{15}


diagram

Figure 1.2.3.1: Two examples of specific effects in transition metal complexes. a) Electron-pair and H-bond donation to an ammine complex from the solvent, and b) a solvent molecule accepting electron density from a cyanide complex.

Acceptor Numbers for common solvents are given below in Table 1.2.3.1.
<table>
<thead>
<tr>
<th>Solvent</th>
<th>Acetone</th>
<th>Pyridine</th>
<th>(CH$_3$)$_2$SO</th>
<th>CH$_3$CN</th>
<th>CH$_2$Cl$_2$</th>
<th>CH$_3$OH</th>
<th>H$_2$O</th>
</tr>
</thead>
<tbody>
<tr>
<td>Acceptor Number</td>
<td>12.5</td>
<td>14.2</td>
<td>19.3</td>
<td>19.3</td>
<td>20.4</td>
<td>41.3</td>
<td>54.8</td>
</tr>
</tbody>
</table>

Table 1.2.3.1: Gutmann Acceptor Numbers for selected solvents. The higher the Acceptor Number, the greater the solvent’s ability to accept electron density from the ligand sphere.$^{15}$

1.2.4 The Influence of Solvent on the Electronic Structure of Transition Metal Cyanide Complexes

The donor-acceptor interactions between transition metal cyanide complexes and solvent molecules perturb the electronic structure of the complexes, giving rise to solvent-dependence in the positions and profiles of their UV/visible absorption spectra. As the Acceptor Number of a solution increases, more electron density is donated from σ(CN) orbitals which are externally directed towards the solvent sphere into low-lying acceptor orbitals on the solvent (Figure 1.2.3.1b). The cyanide ligands become more able to accept π-electron density from the metal centre into π*(CN) orbitals, stabilising dπ(M) orbitals by enhancing dπ(M)-π*(CN) mixing.$^{15}$ This negative solvatochromism increases both the oxidation potential of the metal centre and transition energies between metal-centred orbitals and orbitals centred on spectator ligands (e.g. MLCT transition energies).

The effect is less pronounced in the emission spectra of cyanide complexes. In a $^3$MLCT excited state partial oxidation of the metal centre leads to electron density being shifted from ‘CN’ to ‘M’+, decreasing the available electron density for donor-acceptor interactions ($§1.3$). This decreases the extent of dπ(M)-π*(CN) mixing which leads to the metal centred orbital energies showing less solvent-dependency.$^{15}$

1.3 Spectroscopic and Electrochemical Methods Used in this Work

The analytical techniques employed in this Thesis are UV/vis absorption and luminescence spectroscopy, Fourier-transform infrared (FTIR), and time-resolved infrared (TRIR) spectroscopy and electrochemistry. These provide complementary
information on the electronic structure of the ground and excited states of transition metal complexes.

UV/visible absorption spectroscopy is a simple but powerful spectroscopic technique which provides information on the Franck-Condon excited states populated immediately following the absorption of a photon. As the timescale for photon absorption is very short compared to the timescales for molecular motion, the spectra obtained are representative of the average solvent shell surrounding the molecule prior to absorption (§1.2.1). The shape, intensity and solvent sensitivity of these bands provide information on the relative energies of transitions in metal complexes. The absorption profile obtained from UV/vis spectroscopy is crucial for determining the wavelength of $^1$MLCT transitions for luminescence and TRIR spectroscopy. Insight into other electronic processes e.g. proximity of IL and MC states to MLCT transitions, and presence of IVCT transitions or multiple MLCT bands can also be gained.

Luminescence spectroscopy probes the photons emitted by the complex following excitation. The emissive electronic states are usually the lowest energy excited states in the system and the orbitals involved are usually the HOMO and LUMO. Both the energy and profile of the spectrum, and the lifetime of the emission depend on the electronic structure of the complex and its surrounding environment. The timescale of photon emission is longer than photon absorption (typically nanoseconds or microseconds for metal complexes), and the solvent shell has the opportunity to reorganise itself to solvate the excited state's dipole (§1.2.2, Figure 1.2.2.1). Excited state lifetimes of metal complexes usually follow the 'energy-gap law'. This is a rule-of-thumb which states, for similar complexes, as the energy of the emission decreases, the lifetime also decreases due to an increase in the rate of non-radiative deactivation of the emissive excited state.

Time-Correlated Single Photon Counting (TCSPC) is used to determine the lifetime or time-dependency of emission at particular excitation and emission wavelengths. This technique is very useful for probing complex time dynamics and for comparison with TRIR kinetic data. TCSPC data is often collected using a
'line-of-sight' type apparatus which measures the time elapsed between the excitation event (e.g. laser or lamp flash) and an emitted photon (of a selected wavelength) hitting the detector PMT.

Fourier-transform infrared spectroscopy (FTIR) is a standard analytical technique which enables spectra across the entire mid-infrared range (ca. 50 to 5000 cm\(^{-1}\)) to be measured rapidly. It utilises interferometry to generate interference patterns, which may be processed computationally using Fourier analysis to generate FTIR spectra very quickly (typically less than a second per scan). The information on bond vibrations obtained by FTIR spectroscopy is extremely useful for transition metal complexes containing CO and CN ligands. \(v(\text{CO})\) and \(v(\text{CN})\) bands have high oscillator strengths and narrow, sharp features in a part of the mid-infrared spectrum which is relatively unoccupied by other chemical groups (2100 – 1850 cm\(^{-1}\)).

Time-resolved infrared spectroscopy (TRIR) is a ‘pump-probe’ technique employing a combination of laser excitation and fast infrared detection. It allows the IR spectrum of metal excited states to be recorded down to the picosecond timescale.\(^{17}\) TRIR is particularly useful for the determination of transient molecular structure through the changes in vibrational band positions and profiles through the excitation process. The use of IR ‘reporter’ groups (such as CO and CN) with high oscillator strengths gives intense spectroscopic features, which are sensitive to the electron density of their surroundings. Information about the kinetics of transient formation and decay can also be obtained from TRIR spectroscopy. Further details on the PIRATE apparatus used for ps and ns-TRIR experiments in this Thesis are provided in Chapter 6.

There are some considerable advantages to using TRIR spectroscopy over luminescence spectroscopy for excited state studies of coordination compounds. These advantages are exploited throughout the remainder of this Thesis. TRIR is able to probe ‘dark’ (non-luminescent) excited states in luminescent complexes, as it is not reliant on the detection of emitted photons. For the same reason it may be used to monitor photophysical processes in real time. These attributes are used in Chapter 3 in particular. The study of excited states with very weak or absent
luminescence may also be performed and TRIR can offer insight into the extent of excited state localisation or delocalisation on the IR timescale (< 1 ps). These features are exploited in Chapter 4.

TRIR spectroscopy has been used to study the common classes of excited states of coordination compounds. The general findings are discussed briefly below in reference to IR reporters which are directly ligated to metal centres (e.g. CO and CN ligands). IR reporters which are not directly bonded to metal centres (e.g. amide-substituted bipyridine ligands) behave somewhat differently and are discussed in §1.4.1.1. Upon excitation the change in electronic structure of the metal centre will directly affect the v(CO) or v(CN) bands. TRIR spectra show the difference in absorption between the excited state and ground state spectra. Thus excited state bands are recorded as positive signals and depletions of ground state bands are shown as negative features (Figure 1.3.1).

![Figure 1.3.1: a) The FTIR spectrum of fac-[ReCI(CO)₃(bpy)] in n-PrCN/n-EtCN (5:4 v/v); b) the TRIR spectrum of this solution at 135 K obtained 100 ns after 355 nm excitation. Bands below the axis represent bleaches from the ground state complex and bands above the axis represent new features arising from the excited state species. Adapted from Ref. 20]

1-12
Following the formation of an excited state which decreases the electron density at
the metal centre (e.g. a $^3$MLCT excited state) there is a shift of the $\nu$(CO) or $\nu$(CN) transient bands to higher energy compared to the ground state bands. Conversely
the formation of an excited state which increases the electron density at the metal
centre (e.g. a LMCT, ligand to metal charge transfer excited state) will shift $\nu$(CO)
or $\nu$(CN) bands to lower energy compared to the ground state bands.

The extent of the shifts gives information on the degree of charge transfer in the
excited state as well as indications of the energy gap between ground and excited
states. This may help to identify the nature and location of the excited state in
multichromophoric systems.

In general photoprocesses which do not change the electronic structure of the
metal centre (such as the formation of ligand-centred excited states), lead to no
change in the TRIR spectrum. However in the case of a $\pi,\pi^*$ IL excited state,
where the population of the ligand antibonding orbitals is slightly electron
donating relative to the ground state, a small shift of the excited state $\nu$(CO) bands
to lower energy is observed. This shift is small because the effect is considered to
be of a secondary nature.

Several techniques are commonly used for the probing of electronic structure of
electrochemically generated species. Cyclic voltammetry (CV) determines the
potentials and thermodynamic reversibility of oxidation and reduction processes in
the complex. This provides information on the absolute energies of the orbitals as
opposed to the relative ordering and transition energies from absorption and
emission spectroscopy. As oxidation removes electrons from the HOMO and
reduction adds electrons to the LUMO, photophysics and electrochemistry can be
used to probe the same states in different ways (vide infra).

Spectroelectrochemistry is the spectroscopic study of electrochemically generated
species. The spectra of electrochemically generated compounds can be useful for
comparison with spectra of transients produced through photoexcitation.
UV/visible absorption and IR spectroelectrochemistry are two methods used by
the supramolecular photochemistry community. Insights into the destination of
added electrons, and structural changes in complexes as a result of redox processes allow the comparison of the ‘redox’ electronic structure with the ‘spectroscopic’ electronic structure. Koopmans’ theorem treats the atomic or molecular system as a series of independent one-electron orbitals. The 1st ionisation energy is equivalent to the energy of the HOMO, and the results obtained from spectroscopy and electrochemistry may be the same. Koopmans’ theorem becomes less valid when there are significant changes in the electronic structure of the system upon adding or removing electrons, and when the difference in the timescale of optical and electrochemical processes (i.e. voltammetric vs. Franck-Condon timescales) has a significant impact on the electronic structure and properties of the system.

1.4 Previous Studies of Relevant Complexes and Supramolecular Systems

1.4.1 Monometallic Units

The modular approach to construction of the supramolecular systems in this Thesis is based on the use of the well-known complexes [Ru(bpy)₃]²⁺, [Ru(CN)₄(bpy)]²⁻ and [ReCl(CO)₃(bpy)], where bpy = 2,2'-bipyridine. A brief review of previous research into the spectroscopy, electrochemistry and photophysics of the widely studied [Ru(bpy)₃]²⁺ and [ReCl(CO)₃(bpy)] complexes is presented here. [Ru(CN)₄(bpy)]²⁻ has been less widely studied previously and is discussed at length in Chapter 2.

1.4.1.1 [Ru(bpy)₃]²⁺ and related complexes

The photophysics, electrochemistry and spectroscopy of [Ru(bpy)₃]²⁺ and derivatives of it have been widely studied. Its chemical, photochemical and electrochemical stability and its reactive, energetic, emissive and long-lived 3MLCT excited state make it a good photosensitiser. It is easily synthesised and its
electronic structure can be manipulated to a significant extent through modification of the bipyridyl ligands.\(^23\)

Most mononuclear complexes incorporating the \([\text{Ru(bpy)}_3]^{2+}\) motif have redox and spectroscopic properties similar to the parent complex. The absorption spectrum of \([\text{Ru(bpy)}_3]^{2+}\) in CH\(_3\)CN has intense bands arising from IL \(\pi-\pi^*\) transitions at 185 and 285 nm. Bands at 240 and 450 nm are due to \(1^\text{MLCT}\) transitions, and shoulders at 322 and 344 nm are assigned to metal-centred (MC) transitions. The MLCT bands show weak positive solvatochromism, with the absorption wavelengths of the long wavelength band ranging between 450 and 455 nm.\(^22\) The addition of electron-withdrawing substituents such as amides and esters at the 4,4' positions on the substituted bipyridyl rings, leads to the MLCT bands being observed at longer wavelengths. These ligands are the ultimate electron-acceptors in MLCT transitions.\(^24\) This is due to the stabilisation of the ligand centred \(\pi^*\) orbitals by the substituents.\(^25\) Absorption wavelengths for the low energy transitions of complexes relevant to this work are shown in Table 1.4.1.1.1.

<table>
<thead>
<tr>
<th>Complex</th>
<th>Solvent</th>
<th>(\lambda_{\text{abs}} / \text{nm})</th>
<th>(\lambda_{\text{abs}} / \text{nm})</th>
<th>(\lambda_{\text{abs}} / \text{nm})</th>
</tr>
</thead>
<tbody>
<tr>
<td>([\text{Ru(bpy)}_3]^{2+})</td>
<td>CH(_3)CN</td>
<td>344</td>
<td>415</td>
<td>450</td>
</tr>
<tr>
<td>([\text{Ru(4,4'-(CO}_\text{2Et})_2\text{bpy)}_3]^{2+})</td>
<td>DMF</td>
<td>362</td>
<td>442</td>
<td>471</td>
</tr>
<tr>
<td>([\text{Ru(4,4'-(CONEt_2)_2bpy)}_3]^{2+})</td>
<td>DMF</td>
<td>-</td>
<td>435</td>
<td>459</td>
</tr>
</tbody>
</table>

Table 1.4.1.1.1: Absorption wavelengths for low-energy transitions of \([\text{Ru(bpy)}_3]\)-type complexes relevant to this work. Adapted from Refs.\(^{20,21}\)

The emission spectrum of \([\text{Ru(bpy)}_3]^{2+}\) in CH\(_3\)CN solution at room temperature has components at 615 nm and 650 nm, with an emission lifetime of ca. 850 ns, and a quantum yield of ca. 0.07.\(^23\) Emission occurs from a manifold of three closely lying \(3^\text{MLCT}\) states (10 and 61 cm\(^{-1}\) apart from each other) in thermal equilibrium with each other, which may be effectively treated as a single state. The solvent-dependence of the emission energies, lifetimes and quantum yields of \([\text{Ru(bpy)}_3]^{2+}\) have been investigated at room temperature.\(^17\) Emission wavelengths were found to vary between 605 nm (CH\(_2\)Cl\(_2\)) and 625 nm (H\(_2\)O). The dependence of emission wavelength, emission lifetimes and quantum yields (Table 1.4.1.1.2) were found to
be in accordance with the energy-gap law (§1.3). The emission properties of complexes relevant to this work are shown in Table 1.4.1.1.2.

<table>
<thead>
<tr>
<th>Complex</th>
<th>Solvent</th>
<th>( \lambda_{em} / \text{nm} )</th>
<th>Lifetime / ns</th>
<th>Quantum Yield</th>
</tr>
</thead>
<tbody>
<tr>
<td>[Ru(bpy)]_{3}^{2+}</td>
<td>CH(_{3})CN</td>
<td>615</td>
<td>855</td>
<td>0.07</td>
</tr>
<tr>
<td>[Ru(bpy)]_{3}^{2+}</td>
<td>H(_{2})O</td>
<td>625</td>
<td>630</td>
<td>0.042</td>
</tr>
<tr>
<td>[Ru(4,4'-(CO(_{2})Et)<em>2bpy)]</em>{3}^{2+}</td>
<td>CH(_{3})CN</td>
<td>607</td>
<td>2230</td>
<td>0.3</td>
</tr>
</tbody>
</table>

_Table 1.4.1.1.2: Emission properties of complexes relevant to this work. Adapted from Refs.\(^{17,20,21}\)_

Cyclic voltammetry shows one reversible oxidation for the Ru\(^{II/III}\) couple and a series of bpy-based reductions. The amount, reduction potential and reversibility of these reductions depends on the ligand identity and experimental conditions (e.g. temperature). Several reductions have been observed for complexes with electron-withdrawing substituents (e.g. esters and amides) on the bpy ligands.\(^{25}\) These substituents stabilise low formal oxidation states in the complexes. Table 1.4.1.1.3 shows redox potentials for [Ru(bpy)]\(_{3}\)^{2+} complexes relevant to this work.

<table>
<thead>
<tr>
<th>Complex</th>
<th>Solvent</th>
<th>( E_{%\text{ox}} / \text{V} )</th>
<th>( E_{%\text{red1}} / \text{V} )</th>
<th>( E_{%\text{red2}} / \text{V} )</th>
<th>( E_{%\text{red3}} / \text{V} )</th>
</tr>
</thead>
<tbody>
<tr>
<td>[Ru(bpy)]_{3}^{2+}</td>
<td>CH(_{3})CN</td>
<td>+0.90</td>
<td>-1.72</td>
<td>-1.93</td>
<td>-2.15</td>
</tr>
<tr>
<td>[Ru(4,4'-((CO(_{2})Et)<em>2bpy)]</em>{3}^{2+}</td>
<td>DMF</td>
<td>+1.01</td>
<td>-1.31</td>
<td>-1.43</td>
<td>-1.61</td>
</tr>
<tr>
<td>[Ru(4,4'-((CONEt)<em>2bpy)]</em>{3}^{2+}</td>
<td>DMF</td>
<td>+0.93</td>
<td>-1.54</td>
<td>-1.68</td>
<td>-1.85</td>
</tr>
</tbody>
</table>

_Table 1.4.1.1.3: Redox potentials of complexes relevant to this work. All potentials are quoted vs. Fc/Fc\(^{+}\). Adapted from Refs.\(^{23,25}\)_

A number of TRIR studies of Ru(II) polypyridine complexes have been conducted.\(^{26}\) Ester and amide v(CO) bands of ligands such as (4,4'-((CO\(_{2}\)Et)_2bpy) and (4,4'-((CONEt)_2bpy) provide a spectroscopic probe of the electronic structure of the bpy system. The magnitude of the excited state peak shift carries important information on the extent of charge transfer, the HOMO-LUMO energy gap and the degree of localisation of the excited state, as observed on the IR timescale.
TRIR spectra are also informative on the identity of the ligand accepting the electron density in an MLCT excited state. Information on the energy gap between ground and lowest lying excited state energies has also been extracted from TRIR spectra (§1.4.1.3).

Figure 1.4.1.1.1 shows the excited state spectra obtained for the complexes, [Ru(bpy)$_2$(4,4'-(CO$_2$Et)$_2$tpy)]$^{2+}$, [Ru(bpy)(4,4'-(CO$_2$Et)$_2$tpy)$_2$]$^{2+}$ and [Ru(bpy)$_2$(4,4'-(CONEt)$_2$tpy)]$^{2+}$ in CH$_3$CN. In each of these complexes the lowest energy excited state is (Ru$^{III}$)((4,4'-(COX)$_2$tpy)$^*$. The changes in the spectral profile of the band holds information about the electronic structure of the excited states.

![Figure 1.4.1.1.1: Difference TRIR spectra obtained 0 – 100 ns after 355 nm excitation; a) [Ru(bpy)$_2$(4,4'-(CO$_2$Et)$_2$tpy)]$^{2+}$, b) [Ru(bpy)(4,4'-(CO$_2$Et)$_2$tpy)$_2$]$^{2+}$ and c) [Ru(bpy)$_2$(4,4'-(CONEt)$_2$tpy)]$^{2+}$ in CH$_3$CN. A negative band indicates a decrease in absorbance and a positive band indicates an increase. Adapted from Ref. 24.](image)

The excited state shifts of the spectra in Figure 1.4.1.1.1 a) and b) are identical. As a) contains one (4,4'-(CO$_2$Et)$_2$tpy) ligand and b) contains two, this is consistent with the excited electron being localised on one of the (4,4'-(CO$_2$Et)$_2$tpy) ligands. If it were delocalised, the excited state shift in b) would only be half as
large as a) as the additional electron density would be distributed over both (4,4'-(CO$_2$Et)$_2$bp) ligands. The excited state shift in Figure 1.4.1.1.1 b) is approximately twice that of c). The two complexes are both of the type [Ru(bp$_3$)$_2$(4,4'-(COX)$_2$bp)$_2$]$^{2+}$, with a) X=OEt and b) X=NEt$_2$. The smaller excited state shift in the amide-substituted complex indicates that there is a smaller change in electron density in the excited state of c) than in a). This is consistent with a greater extent of charge transfer in [Ru(bp$_3$)$_2$(4,4'-(CO$_2$Et)$_2$bp)$_2$]$^{2+}$ than [Ru(bp$_3$)$_2$(4,4'-(CONEt$_2$)$_2$bp)$_2$]$^{2+}$ as the amides are less electron-withdrawing than the esters.

1.4.1.2 fac-[ReCl(CO)$_3$(bp)]

N. B.: In this Thesis all rhenium (I) tricarbonyl complexes referred to will be facial isomers and the "fac" notation may not be explicitly stated but is implied in all cases.

![Fac-Complex Diagram]

The photophysics, electrochemistry and spectroscopy of fac-[ReCl(CO)$_3$(bp)] and derivatives of it have been studied extensively in the literature, albeit to a lesser extent that polypyridyl complexes of ruthenium (II). In a similar way to [Ru(bp$_3$)$_3$]$^{2+}$, rhenium (I) complexes of bp also exhibit relative chemical, photochemical and electrochemical stability, although the carbonyl groups are labile under certain conditions. [ReCl(CO)$_3$(bp)] has a reactive, energetic, emissive and long-lived 3MLCT excited state making it a good photosensitiser. It is easily synthesised and its electronic structure can be manipulated to a significant extent through modification of the pyridyl ligands. Key differences with ruthenium tris-bipyridyls is the option of adding further ancillary ligands through substitution of carbonyl groups and the visibility of CO groups in IR and Raman vibrational spectra. This means that [ReCl(CO)$_3$(bp)]-based systems require no further modification to act as spectroscopic reporters. To this end it is a key advantage of this system over other common transition metal complexes that the complex possesses carbonyl ligands which are directly ligated to the metal centre and are
able to act as sensitive reporters of electronic density. This effect is exploited in Chapters 3 and 4 of this Thesis.

The absorption spectrum of \([\text{ReCl(CO)}_3(\text{bpy})]\) in \(\text{CH}_3\text{CN}\) has bands arising from IL \(\pi-\pi^*\) transitions below and around 300 nm. The lowest energy absorption maximum is in the visible region of the spectrum at 370 nm and arises from a \(d\pi(\text{Re}) \rightarrow \pi^*(\text{bpy})\) \(^1\text{MLCT}\) transition. The \(^1\text{MLCT}\) band shows weak positive solvatochromism, with the absorption wavelengths of the long wavelength absorption ranging between 370 and 384 nm in \(\text{CH}_2\text{Cl}_2\).\(^{29}\) The ultimate electron-acceptor in MLCT transitions of \([\text{ReCl(CO)}_3(\text{bpy})]\) is the sole polypyridyl ligand and there is no debate over excited state locus and timescale of localisation due to the lack of multiple polypyridyl ligands. This is in contrast to scientific investigations of \([\text{Ru(bpy)}_3]^{2+}\) (\textit{vide supra}).

The emission spectrum of \([\text{ReCl(CO)}_3(\text{bpy})]\) in \(\text{CH}_3\text{CN}\) solution at room temperature has a maximum of 612 nm with an emission lifetime of \(\text{ca. } 25\) ns and a quantum yield of \(\text{ca. } 0.06\).\(^{29,30}\) The broad, featureless emission profile is indicative of \(^3\text{MLCT}\) emission from the energetically lowest-lying \(d\pi(\text{Re}) \rightarrow \pi^*(\text{bpy})\) excited state. The solvent-dependence of the emission energies and lifetimes of \([\text{ReCl(CO)}_3(\text{bpy})]\) have been investigated at room temperature, although the sparing solubility of the complex prevents its study in most polar solvents. Emission wavelengths were found to be essentially invariant in \(\text{CH}_2\text{Cl}_2\) and \(\text{CH}_3\text{CN}\), although lifetimes were found to be longer in \(\text{CH}_3\text{CN}\) (50 ns) than in \(\text{CH}_2\text{Cl}_2\) (25 ns).\(^{29}\)

Cyclic voltammetry of \([\text{ReCl(CO)}_3(\text{bpy})]\) in \(\text{CH}_3\text{CN}\) shows one irreversible oxidation (\(E_p = +1.32\) V (vs SCE)) for the \(\text{Re}^\text{VII}\) couple, and one reversible bpy-based reduction (\(E_{1/2} = -1.35\) V vs SCE).\(^{29}\)

A number of investigations have utilised the IR reporting properties of \([\text{ReCl(CO)}_3(\text{bpy})]\) between 1850 and 2100 cm\(^{-1}\) and relatively long excited state lifetimes to study its ground state properties using FTIR and its excited state properties using TRIR spectroscopy (see Figure 1.3.1).\(^{30,31}\) In \(\text{CH}_3\text{CN}\), the ground state complex has FTIR bands at 2023, 1917 and 1899 cm\(^{-1}\). The IR band profile
results from the $C_3$ symmetry of the complex, which gives rise to three infrared-active $\nu$(CO) modes $[A'(1) + A'(2) + A'\prime']$. Upon formation of the $^3$MLCT excited state the parent bands are bleached and three new bands are formed at 2069, 1990 and 1957 cm$^{-1}$.\textsuperscript{32} These results are closely comparable with those measured in CH$_2$Cl$_2$,\textsuperscript{33} in keeping with the similarity of other spectroscopic and photophysical properties of [ReCl(CO)$_3$(bpy)] in low A. N. solvents. Excited state shifts for rhenium (I) tricarbonyl complexes were later shown by both theoretical and experimental approaches (see §1.4.1.3) to be non-trivial, with the ordering of ground and excited state IR bands being significantly different \textit{(vide infra)}.

1.4.1.3 Other Monometallic Complexes

An interesting extension to conventional pump-probe spectroscopy of rhenium (I) complexes has recently been demonstrated by Bredenbeck \textit{et al.} in a study of [ReCl(CO)$_3$(4,4'-(CH$_3$)$_2$-bpy)].\textsuperscript{34} Using three laser pulses (one to excite the molecule into an electronically excited state and the other two to label/excite and probe vibrational states), they have performed both labelling and hole-burning 't2D-IR' experiments which have afforded considerable insight into the nature of the bands present in the ground and excited state infrared spectra of rhenium (I) tricarbonyl complexes (Figure 1.4.1.3.1).

Their findings reinforced previous experimental and theoretical work carried out by Dattelbaum \textit{et al.} \textit{(vide infra)} which suggested that the ordering of ground and excited state bands in the infrared spectra were not straightforward. Some bands were shifted to a much greater extent than others, with $\Delta\nu_{\text{CO}}(\text{Excited State}) A'(1) > A'(2) > A''$ \textit{(vide infra)}.\textsuperscript{35} The study by Dattelbaum \textit{et al.} relating ground-to-excited state energy gap to $\Delta\nu_{\text{CO}}$ shift in the complexes examined on the complexes fac-[Re(4,4'-X$_2$-bpy)(CO)$_3$(4-EtPy)]$^+$ (where X is H, CH$_3$, OCH$_3$, or CO$_2$Et; 4-EtPy is 4-ethylpyridine). Systematic variations in $\Delta\nu_{\text{CO}}$ were observed with the excited-to-ground state energy gap, derived from a Franck-Condon analysis of emission spectra. The energy gap was tuned using the X$_2$ substituents on the 4,4'-substituted bpy ligand. It was found that variations in $\nu_{\text{CO}}$(Ground State) due to differences in backbonding in the complexes are small. A pronounced trend in $\nu_{\text{CO}}$(Excited State) was observed, as the $\nu_{\text{CO}}$(Excited State) wavenumber decreased with
increasing energy gap. The mode specific manner of the variations in ground-to-excited state band shifts were also evaluated (vide supra), with $\Delta v_{\text{CO(ES-GS)}} A'(1) > A'(2) > A''$.\textsuperscript{35}

![Graph](image)

**Figure 1.4.1.3.1: Top:** a) FTIR and b) 1-D ps-TRIR spectrum (at 20 ps after excitation) of [ReCl(CO)$_3$((CH$_3$)$_2$-bpy)] with previous (red) and new (green) excited state assignments. **Bottom:** T2D-IR spectrum (at 2 ps after excitation) of [ReCl(CO)$_3$((CH$_3$)$_2$-bpy)] with cross-correlation peaks identified from the two different pulse sequences employed (left and right). Adapted from Ref.\textsuperscript{34}

A recent example of a photophysical investigation into a complex with a solvent switchable excited state type was demonstrated by Pomestchenko et al.\textsuperscript{36} In this study a multichromophoric platinum (II) complex, [Pt(dbppy)(C=C-nap)$_2$], where dbppy = 4,4'-{(di-tert-butyl)-2,2'-bipyridine and C=C-nap = 1-ethynylnapthalene, was studied in four different solvent environments. As the $d\pi$(Pt) $\rightarrow \pi^*(dbppy)$
MLCT transitions in this complex exhibited negative solvatochromism, increasing solvent polarity raised the energy of the $^3$MLCT state.

In low polarity solvents, the emission spectra obtained are consistent with emission from the lowest energy $^3$MLCT excited state. However, in higher polarity solvent environments the spectra were significantly different, with the close proximity in energy between $^3$MLCT and $^3$IL states leading to different photophysical behaviour. It was postulated that rather than being a simple 'switch over' of the nature and locations of the lowest energy excited state, there were significant excited state mixing or equilibrium effects observed. The authors found it difficult to be conclusive regarding the excited state structures observed. This was perhaps because the optical techniques used (time-resolved absorption and emission spectroscopies) give less detailed information about excited state structure than a combined study also incorporating transient vibrational spectroscopic techniques such as TRIR.

1.4.2 Photoinduced Processes in Polynuclear Systems

There have been numerous spectroscopic studies of polynuclear systems but here we shall focus on a small number of systems relevant to this Thesis in which photoinduced processes are studied using spectroscopic approaches similar to those employed in this Thesis.

1.4.2.1 Polypyridyl-linked systems

An interesting study of PEnT in a bimetallic Ru(II)/Re(I) system bridged by a non-symmetric polypyridyl ligand with two chemically different binding sites was reported by Schoonover and co-workers. The isomeric complexes $[(\text{bpy})_2\text{Ru-AB-Re(CO)}_3\text{Cl}]$ (RuABRe) and $[(\text{OC})_3\text{ClRe-AB-Ru(bpy)}_2]$ (ReABRu, where AB = 2,2':3',2'':6'',2'''-quaterpyridine) were prepared and studied by luminescence and TRIR spectroscopy. The two binding sites, A and B are inequivalent with B being more sterically hindered (Figure 1.4.2.1.1).
From an analysis of emission lifetimes it had been previously demonstrated that at room temperature the emission observed in the RuABRe isomer was Ru-based and for ReABRu it was Re-based. However this result was not definitive as it relied on emission lifetime measurements which may have been influenced by an excited state equilibrium in the ReABRu case, raising the possibility that the Ru-based excited state was lower in both cases. With the emission results proving to be ambiguous, TRIR spectroscopy was employed to probe the excited states involved in PEnT in more detail.

For the complex RuABRe, ps and ns-TRIR spectroscopy gave equivalent spectra, with v(CO) excited state bands at 1899, 1921 and 2023 cm⁻¹. These were shifted to lower wavenumber than the FTIR bands recorded for the ground state complex at 1888, 1910 and 2016 cm⁻¹. The shift of the Re-based v(CO) bands to lower wavenumber is consistent with the formation of a Ru-based ³MLCT excited state, as the rhenium centre is reduced by the additional electron density on the bridging ligand. The excited state lifetime was measured to be ca. 400 ns.³⁸

The isomer ReABRu showed differences in its ps and ns-TRIR spectra. The TRIR spectrum at 100 ps contained ground-state bleaches at 1902, 1922 and 2028 cm⁻¹ and excited state bands at 1892, 1912 and 2018 cm⁻¹ but also revealed the presence of a broad positive feature near 1965 cm⁻¹ (see Figure 1.4.2.1.1). This is consistent with both Re-based and Ru-based ³MLCT excited states being present at early times - with a statistical mixture of the two transient species present in the ps-TRIR spectrum. Interestingly luminescence spectroscopy only detected the presence of a Re-based ³MLCT excited state as the emission lifetime (ca. 20 ns) is

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Figure 1.4.2.1.1: The ligand AB.

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³⁸
determined largely by the non-radiative decay rate of the Re-based transient. This highlights the 'invisibility' of some photophysical processes and excited states to luminescence spectroscopy.\textsuperscript{38}

This case study illustrates the utility of TRIR in the probing and assigning of excited states and photophysical processes, even with only one of the two metal centres present possessing an IR active reporter group. The polynuclear complexes in this Thesis are designed with reporter groups on both metal centres, leading to a potentially richer data set when interpreting and analysing the photophysical processes occurring.

![Figure 1.4.2.1.1: TRIR difference spectra of RuABRe (solid line) and ReABRu (dotted line) 100 ps after 405 nm excitation in CH\textsubscript{3}CN. Reproduced from Ref.38](image)

Whilst not directly related to the systems described in this Thesis, the reader is directed to an excellent review of the use of ruthenium (II) bipyridine complexes in supramolecular systems by Balzani \textit{et al.}, which highlights the current state of the art in supramolecular machine and device preparation. This perhaps represents the level of intermolecular interaction which can be aimed for in the future with the prospective device components contained in this Thesis.\textsuperscript{39}
1.4.2.2 CN linked systems

An interesting use of emission, transient absorption and TRIR spectroscopies to study energy transfer in bimetallic ruthenium (II) and rhenium (I) complexes was presented by Schoonover *et al.* and Thompson *et al.* over the course of two publications.\(^{40,41}\) In their study, ostensibly the first using ps and ns-TRIR to probe intermetallic PEnT, the metal centres were linked using a cyanide bridge. Due to the asymmetry of the cyanide linkage two isomers are possible and the study focussed on comparing the photophysical properties of these two linkage isomers. Previous studies found that energy transfer characteristics in a homobimetallic complex \([(NC)(bpy)_2Ru(CN)Ru(bpy)_2(CN)]\) were strongly influenced by the redox asymmetry induced by the bridge.\(^{42}\) In both complexes \([(phen)(OC)_3Re(CN)Ru(bpy)_2(CN)]^+\) and \([(phen)(OC)_3Re(NC)Ru(bpy)_2(CN)]^+\) efficient Re$\rightarrow$Ru PEnT was found to occur. Multiple excited state species were generated and this was attributed to the unselective excitation of either terminus upon absorption of 355 nm laser light. Unfortunately the study did not include a complete TRIR dataset for both linkage isomers and is not described in detail here.

There have also been a few other noteworthy studies using TRIR to study electron transfer in cyanide-bridged mixed valence bimetallic complexes of d$^6$ metals such as \([(NC)_5Fe^{III}(CN)Ru^{II}(NH_3)_3]^+\) and \([(NC)_5Os^{II}(CN)Ru^{III}(NH_3)_3]^+\).\(^{43,44}\) In these studies, optical excitation incited forward electron transfer upon population of Franck-Condon excited states, with ultrafast TRIR used to monitor the back electron-transfer process. The focus of these studies by Gilbert and co-workers was to probe the environment-solute coupling through analysis of the spectral dynamics, which is outside of the scope of this work.

1.5 Closing Points

This Chapter has described some of the basic theory supporting the content of later Chapters of this Thesis. The key concepts to take forwards are listed below.

- Supramolecular photophysics may involve the interaction of a potentially unlimited number of components comprised of coordination compounds.
• Transition metal complexes with polypyridyl ligands often exhibit long-lived triplet MLCT excited states which may be utilised as the first step in a supramolecular photoprocess.

• Optical spectroscopic techniques such as UV/vis absorption and luminescence spectroscopy, probe molecular electronic structure and its interaction with the solvent environment on different timescales.

• The prevalent mechanism of photoinduced interaction in a supramolecular system may be influenced by the choice of chromophores and design of the linkage between them.

• Förster-type PEnT is extremely sensitive to the distance between and relative orientation of donor and acceptor chromophores.

• Solvatochromism may have a major effect on the energy levels of transition metal complexes.

• The ability of a solvent to exhibit SSDA interactions may be quantified using the Gutmann Acceptor Number theory.

• Metal-cyanide complexes display remarkably strong negative solvatochromism due to specific solvent-solute interactions.

• Some ligands or chemical groups can act as efficient reporters of electronic structure information in TRIR spectroscopic experiments, potentially enabling the monitoring of PEnT and/or PET in situ.

• Having access to multiple spectroscopic techniques and/or reporter groups increases the amount of spectral information available and allows a more complete understanding of the systems under study.

• Luminescence spectroscopy is a powerful tool when the emitted photon carries desired information about the photophysics of the system being studied. However the ‘invisibility’ some photoprocesses to luminescence - ‘dark’, non-radiative or photonless resonance processes – renders the technique less suitable for their study.

• TRIR spectroscopy is a powerful tool for elucidating electronic structure of excited states when ‘vibrational reporter’ chemical groups are present.

• The combined use of luminescence and TRIR spectroscopies may reinforce the findings of each other and reveal subtleties that either technique may have not have raised.
The complexes [Ru(bpy)$_3$]$^{2+}$ and fac-[ReCl(CO)$_3$(bpy)] have well-studied and characterised excited states which are amenable to incorporation into supramolecular systems.

Photophysical study of coordination compounds with spectrally similar or multiple chromophores may produce more complicated or convoluted spectral or kinetic information due to unselective excitation producing multiple excited state species.

Transition metal complexes may be incorporated into larger supramolecular devices.

Chapter 2 will present a study of the [Ru(CN)$_6$(NN)$]^{2-}$ (where (NN) = bidentate α-diimine) family of metal complexes. Chapter 3 will present a study of weakly interacting bimetallic supramolecular systems designed to undergo through-space (Förster) PEnT. Chapter 4 will present a study of strongly interacting bimetallic complexes designed to undergo through-bond energy / electron transfer. Chapter 5 will provide an overall summary, discussion and closing points and presents some interesting future directions for research in this area. Chapter 6 will describe the experimental and theoretical approaches used in this Thesis.
1.6 References


Chapter 2:

Tetracyanoruthenate (II) Complexes With $\alpha$-Diimine Ligands
Chapter 2: Tetracyanoruthenate (II) Complexes With u-Diimine Ligands

2.1. Introduction

In this Chapter a series of tetracyanoruthenate (II) complexes with bidentate α-diimine ligands are presented along with the results of electronic and vibrational spectroscopic and electrochemical measurements. The findings are discussed with reference to previously published studies.

2.1.1. [Ru(CN)₄(bpy)]²⁻

[Ru(CN)₄(bpy)]²⁻ was first reported by Bignozzi et al. as a mono-bipyridyl analogue of [Ru(bpy)₃]²⁺. The extent and timescale of delocalisation of the lowest energy 3MLCT excited state of [Ru(bpy)₃]²⁺ was (and still is) the subject of considerable debate. [Ru(bpy)₃]²⁺ has been compared with [Ru(CN)₄(bpy)]²⁻ where two bpy ligands have been replaced by strong field CN ligands. These studies addressed some of the unanswered questions regarding excited state delocalisation in ruthenium (II) polypyridyl complexes. [Ru(CN)₄(bpy)]²⁻ possesses a localised 3MLCT excited state and the strong field cyanide ligands increase the energy difference between metal centred (MC) states thus raising the lowest 3MC excited state energy and reducing the occurrence of radiationless deactivation following thermal promotion to these states.

Several studies have focused on [Ru(CN)₄(bpy)]²⁻ and its behaviour in varying solvent environments. Its spectroscopic, electrochemical and photophysical properties exhibit strong negative solvatochromism due to donor-acceptor (SSDA) interactions between the cyanide ligands and solvent molecules (§1.2.3). Table 2.1.1.1 shows the absorption and emission wavelengths and oxidation potentials for [Ru(CN)₄(bpy)]²⁻ in a number of solvents. λ abs¹ is assigned to an intraligand π-π* transition and λ abs² and λ abs³ are assigned to ¹MLCT transitions. The solvent-dependence of the IL transitions is due to mixing between the dπ(Ru) and π(bpy) orbitals, which introduces some MLCT character into the IL transitions.
Table 2.1.1.1: Absorption and emission wavelengths and oxidation potentials for 
[Ru(CN)₄(bpy)]²⁻ in a selection of solvents. Adapted from Ref.⁴

<table>
<thead>
<tr>
<th>Solvent</th>
<th>A. N.</th>
<th>( \lambda_{\text{abs}}^1 )/ nm</th>
<th>( \lambda_{\text{abs}}^2 )/ nm</th>
<th>( \lambda_{\text{abs}}^3 )/ nm</th>
<th>( \lambda_{\text{em}} )/ nm</th>
<th>( \text{E}_{\text{V/ox}} )/ V</th>
</tr>
</thead>
<tbody>
<tr>
<td>H₂O</td>
<td>54.8</td>
<td>285</td>
<td>300</td>
<td>404</td>
<td>621</td>
<td>+0.77</td>
</tr>
<tr>
<td>CH₃OH</td>
<td>41.3</td>
<td>291</td>
<td>329</td>
<td>450</td>
<td>684</td>
<td>+0.27</td>
</tr>
<tr>
<td>CH₃CH₂OH</td>
<td>37.1</td>
<td>293</td>
<td>342</td>
<td>471</td>
<td>712</td>
<td>+0.11</td>
</tr>
<tr>
<td>CH₂Cl₂</td>
<td>20.4</td>
<td>297</td>
<td>370</td>
<td>528</td>
<td>776</td>
<td>-</td>
</tr>
<tr>
<td>CH₃CN</td>
<td>19.3</td>
<td>295</td>
<td>374</td>
<td>535</td>
<td>805</td>
<td>-0.18</td>
</tr>
</tbody>
</table>

Linear correlations between the absorption and emission energies in ruthenium (II) cyanide complexes and the Acceptor Number of the solvent were found by Timpson et al. and can be explained using the theory of SSDA interactions in cyanide complexes (§1.2.4).⁴

The factors affecting the lifetime and quantum yield of [Ru(CN)₄(bpy)]²⁻ and similar complexes in normal and deuterated solvents have been studied by Kovács and Horváth.⁷,⁸ Lifetimes were found to be around 3 times longer in D₂O (355 ns) and CD₃OD (162 ns) as compared with H₂O (113 ns) and CH₃OH (62 ns) and quantum yields were similarly enhanced approximately threefold. The enhancement of the emission lifetime in deuterated solvents has been attributed to a decrease in the rate of non-radiative deactivation through a fourth \(^3\text{MLCT}\) state which lies ca. 1000 -1400 cm⁻¹ above the ground state manifold of three \(^3\text{MLCT}\) states in thermal equilibrium (Figure 2.1.1.1).

Figure 2.1.1.1: A schematic of the decay pathways in the [Ru(CN)₄(bpy)]²⁻ system. Thermal population of the \(^4\text{MLCT}\) state provides another non-radiative decay pathway for the excited complex. Reproduced from Ref.⁷

2-2
The strong influence of solvent deuteration on the excited state decay rates shows the importance of hydrogen bonding in the relaxation mechanism. This study also included a brief discussion of similar results obtained with the related complexes [Ru(CN)₄((5,5-CH₃)₂bpy)]²⁻, [Ru(CN)₄(phen)]²⁻ and [Ru(CN)₄((C₆H₅)₂-bpy)]²⁻ and a subsequent article presented a temperature-dependence study of the luminescence of [Ru(CN)₄(bpy)]²⁻ (outside the scope of this work) and a visual interpretation of the SSDA interactions taking place (Figure 2.1.1.2). ⁸

![Figure 2.1.1.2: Schematic of the SSDA interactions between [Ru(CN)₄(bpy)]²⁻ and aqueous/alcoholic solvents. Reproduced from Ref.⁸](image)

A TRIR study of a complex related to [Ru(CN)₄(bpy)]²⁻ has been carried out (Figure 2.1.1.3 a).⁹ Although the polypyridyl ligand is somewhat different from bpy, similar behaviour upon excitation would be expected, with a (RuІІІ)(L²⁻) ³MLCT excited state being formed. The difference TRIR spectrum (Figure 2.1.1.3b) in the v(CN) region shows a transient (positive peak) at ca. 50 cm⁻¹ higher than the peak in the ground state spectrum (Figure 2.1.1.3c). The shift to higher wavenumber in the excited state is consistent with the formation of a ³MLCT excited state. Partial oxidation of RuІІ to RuІІІ increases the amount of Ru-C backbonding (Figure 2.1.1.2) which strengthens the C-N bond. This increases the energy of the v(CN) stretch, leading to the band’s increase in wavenumber. The study also demonstrated considerable pH-dependence on the emission properties of [Ru(CN)₄(L₂)]²⁻ which is outside the scope of this work.

Much of the above has recently been discussed in a timely review article.¹⁰
2.1.2. [Ru(CN)₄(bpm)]²⁻

An important property of the 2,2'-bipyrimidine ligand is its ability to chelate two metal centres within 550 pm of each other across a planar, conjugated network thus facilitating electronic coupling between the two units. This effect forms the basis of Chapter 4. [Ru(CN)₄(bpm)]²⁻ has been less widely studied than [Ru(CN)₄(bpy)]²⁻ and only one published study has investigated the solvatochromism of this complex.

Samuels and DeArmond have conducted a brief study of the spectroscopy, photophysics and electrochemistry of [Ru(CN)₄(bpm)]²⁻. The absorption bands corresponding to ¹MLCT transitions show considerable negative solvatochromism in CH₃CH₂OH and DMF (Figure 2.1.2.1a). This trend is consistent with the effects of specific solvent-cyanide interactions (§1.2.3), with λₘₐₓ = 375 and 490 nm in ethanol and λₘₐₓ = 425 and 590 nm in DMF. The low-temperature emission spectrum in CH₃CH₂OH shows λₘₐₓ = 550 nm, with shoulders at 580 and 620 nm (Figure 2.1.2.1b). This ³MLCT emission wavelength is significantly shorter than in
[Ru(CN)₄(bpy)]²⁻, which has \( \lambda_{\text{max}} = 712 \text{ nm} \) in ethanol.⁴ No emission was detected at 298 K. Thus the energies of \(^3\)MLCT emission of [Ru(CN)₄(bpy)]²⁻ and [Ru(CN)₄(bpm)]²⁻ have been shown to be considerably different (vide supra, §2.2.1).

Cyclic voltammetry of [Ru(CN)₄(bpm)]²⁻ showed one reversible reduction at \( E_{1/2} = -2.09 \text{ V} \) (vs. Fe/Fe⁺) in DMF at a less negative potential than the corresponding couple in [Ru(CN)₄(bpy)]²⁻, \( E_{1/2} = -2.46 \text{ V} \).⁴,¹² Oxidation potentials for the Ru\(^{II}/III\) couple were not reported. Thus [Ru(CN)₄(bpm)]²⁻ has lower-lying unoccupied ligand-based orbitals (from electrochemistry) and lower energy MLCT transitions (from spectroscopy) than [Ru(CN)₄(bpy)]²⁻.

![Figure 2.1.2.1: a) Absorption and b) low-temperature emission spectra for [Ru(CN)₄(bpm)]²⁻ in ethanol (solid line) and DMF (dotted line). Adapted from Ref.¹²](image)

2.1.3. [Ru(CN)₂(bpy)₂] and Other Ruthenium Cyanide Complexes

The related complex [Ru(CN)₂(bpy)₂] was first prepared and characterised in the late 1960s by Crosby and co-workers,¹³,¹⁴ who noted and discussed the differences in absorption spectra obtained in CH₃OH and DMF solutions. Later studies compared ruthenium (II) polypyridyl complexes with 1-4 CN groups, showing that the degree of solvatochromism observed was linearly related to the number of CN groups ligated to
the metal centre. The pH and temperature dependence of the luminescence and TRIR spectrum of [Ru(CN)2(bpy)2] in solution have also been reported.

2.1.4. Summary of [Ru(CN)4(NN)]2- Properties

The salient points to be taken from the literature reviewed above may be summarised in terms of three significant properties of the [Ru(CN)4(NN)]2- unit. Firstly, its strong negative solvatochromism which arises from (SSDA) interactions between the externally-directed electron lone pairs on the cyanide ligands and solvent molecules. This means that its 3MLCT energy and its ground and excited state redox potentials are highly solvent dependent, for example the 3MLCT luminescence maximum varies from 640 nm in H2O to 818 nm in (CH3)2NCHO, and the RuIII redox potential varies from +0.77 V to -0.28 V (vs. Fe/Fe+) between the same solvents. This also applies in the solid state as demonstrated by Evju et al. using a crystalline salt of [Ru(CN)4(bpy)]2- as a humidity sensor on the basis of its reversible colour change from purple to yellow in the presence of humidity. This effect arises because atmospheric H2O molecules form hydrogen bonds to the externally directed lone pairs of the cyanide ligands, modulating the MLCT transition energy and as a result the colour of the complex.

Secondly, the IR-active cyanide ligands of [Ru(CN)4(NN)]2- allow the transfer of excitation energy and electrons in transition metal complexes to be monitored by time-resolved infrared (TRIR) spectroscopy, since the v(CN) vibrations are sensitive to redistribution of electron density in the complex. One relevant example of the application of TRIR to a ruthenium (II) cyanide complex (§2.1.1) has provided some background and insight into the spectroscopic behaviour of tetracyanoruthenate (II) polypyridyl complexes.

Finally, the externally directed lone pairs on the nitrogen atoms of the cyanide ligands of [Ru(CN)4(NN)]2- are capable of acting as hydrogen bond acceptors, allowing the complex to act as a component of hydrogen bonded assemblies; and the lone pairs can
also coordinate to additional metal cations, allowing formation of polynuclear complexes via Ru–CN–M linkages (for examples see §1.4.2.2) or the binding of bare metal ions (metallochromism, §3.5.3).\textsuperscript{19,20}

These properties have not been fully exploited previously but form a major part of the work described in this Thesis. We have designed systems to exploit these advantages which present new applications for transition metal polypyridyl complexes in supramolecular photochemistry. Furthermore analogues and derivatives of [Ru(CN)\textsubscript{4}(bpy)]\textsuperscript{2−} have been scarcely studied, beyond addition of methyl substituents or simple replacement of bipyridine by phenanthroline.\textsuperscript{7}

2.2. Scope, Results and Discussion

A range of [Ru(CN)\textsubscript{4}(NN)]\textsuperscript{2−} complexes (shown below) were studied using UV/vis absorption, luminescence, FTIR and TRIR spectroscopy and cyclic voltammetry in D\textsubscript{2}O and CH\textsubscript{3}CN where not limited by solubility or solvent windows.
2.2.1. UV/vis Absorption and Luminescence Spectroscopy

In aqueous solution all of the complexes display an absorption band arising from a Ru \textarrow{\textrightarrow} diimine $^1$MLCT transition at $\lambda_{abs} \sim 400$ nm (Table 2.2.1.1, spectra not shown). RuCNBPBM is unique in this series and exhibits two clearly-resolved transitions in water at 342 and 438 nm. Upon cation exchange (replacing K$^+$ with PPN$^+$) some of the complexes were also soluble in CH$_3$CN; in these cases comparison with the spectra recorded in D$_2$O shows the extent of negative solvatochromism exhibited in complexes of this type. In CH$_3$CN the $^1$MLCT absorption profile is resolved into two components, one of which is at much lower energy than in D$_2$O. For example in RuCNdmb the MLCT absorption in water occurs at 393 nm, whereas in CH$_3$CN there are two bands at 372 and 514 nm. For RuCNBPBM the two transitions apparent in D$_2$O (342, 437 nm) are red-shifted to 415 and 575 nm respectively in CH$_3$CN.

<table>
<thead>
<tr>
<th>Complex</th>
<th>Solvent</th>
<th>$\lambda_{abs,max}$ / nm</th>
<th>$\lambda_{em}$ / nm</th>
<th>$\tau$ / ns</th>
<th>$\Phi$</th>
<th>$E_{RuII/III}$ / V</th>
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<tbody>
<tr>
<td>RuCNbpy</td>
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<td>285, 300, 404</td>
<td>620</td>
<td>360</td>
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<td>D$_2$O</td>
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<td>624</td>
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<td>RuCNdmb*</td>
<td>D$_2$O</td>
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<td>621</td>
<td>370</td>
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<td></td>
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</tbody>
</table>

Table 2.2.1.1: UV/visible and luminescence data for tetracyanoruthenate (II) complexes. $^*$UV/vis and luminescence spectra of RuCNdmb and RuCNBPBM are shown and discussed in detail in later Chapters (§3.2.1 and §4.3.1). CV in H$_2$O with 0.1 M KCl electrolyte. $^#$Irreversible Processes.

---

1 The spectra for RuCNdmb and RuCNBPBM are presented and discussed in detail in Chapters 3 and 4 respectively.
In every case in D$_2$O solution, excitation into the $^1$MLCT absorption transition results in broad, structureless luminescence in the range 550-670 nm which arises from the $^3$MLCT excited state, consistent with the extensive literature studies of ruthenium (II) polypyridyl complexes.$^{21}$ Kovács and Horváth recently showed that using D$_2$O as the solvent affords considerably longer-lived luminescence from complexes of this type than in H$_2$O due to attenuation of non-radiative decay pathways (§2.1.1).$^8$ As discussed above the complexes RuCNdmdb and RuCNphen have been reported previously and will not be discussed in detail, although reinvestigating them gave satisfactory agreement between the results here and previously published studies.$^{17}$

In RuCNtpy the $^1$MLCT absorption and $^3$MLCT luminescence are red-shifted by 35 nm compared to RuCNbpy, despite the donor set being nominally identical. In the solid state, this may be ascribed to the long bond Ru(1)–N(8) (2.19 Å),$^{22}$ involving the central ring of the terpyridyl ligand, which is lengthened by virtue of having the sterically bulky pyridyl residue at its C$^2$ position, ortho to the site of coordination. Lengthening this bond will weaken the ligand field of the coordinated bipyridyl fragment compared to normal un-encumbered bpy (Ru–N separation ca. 2.10 Å), which will slightly raise the d(π) orbital set and hence reduce the MLCT absorption and emission energies. This has been observed previously in derivatives of [Ru(bpy)$_3$]$^{2+}$ in which the bpy ligands have a bulky substituent at the C$^2$/C$^6$ position.$^{23}$

Variation of the nature of the α-diimine ligand results in significant changes to the $^1$MLCT absorption and $^3$MLCT luminescence energies for electronic reasons. The highest-energy absorption and luminescence in the series come from RuCNpypz, implying that the pypz ligand has a higher-energy LUMO than the other diimines. This is corroborated by the short luminescence lifetime (17 ns) which is a consequence of the high-lying $^3$MLCT excited state being close in energy to the metal-centred d-d state, which provides a facile thermally-activated route for radiationless deactivation. Despite the relatively short luminescence lifetime, the high $^3$MLCT energy of RuCNpypz will make it of particular value as an energy-donor in polynuclear assemblies (e.g. for use as a charge-injector into nanocrystalline TiO$_2$).$^{24}$

2-9
The luminescence of RuCNBPM at room temperature is too weak to allow the excited state lifetime to be measured by our apparatus. Replacement of bipyridine by bipyrimidine has been shown in other complexes to result in much weaker, shorter-lived emission (see Chapter 4 for a full discussion of this effect). For those complexes studied in CH$_3$CN, the $^3$MLCT emission maxima (like the $^1$MLCT absorption maxima) are substantially red-shifted ($\lambda_{\text{max}} > 700$ nm) compared to the situation in water, with very weak (often barely detectable) and short lived emission. The luminescence quantum yields in CH$_3$CN are at least 10 times smaller than in water, with short lifetimes (usually $\tau < 5$ ns) that were difficult to measure by luminescence methods because of the weakness of the signal.

2.2.2. Electrochemistry

Cyclic voltammetric studies of the potassium salts of the tetracyanoruthenate (II) complexes in water showed that all but one of the complexes possess a reversible, one-electron oxidation process. By analogy with the known behaviour of [Ru(CN)$_4$(bpy)]$^{2-}$, this couple may be assigned as a metal-centred Ru$^{II/III}$ couple. The redox potentials measured are listed in Table 2.2.1.1. In most cases this process is reversible (peak-peak separation 60-80 mV, equal cathodic and anodic peak currents, peak current proportional to square root of scan rate) on the voltammetric timescale. For RuCNtpy however the process is completely irreversible with no return wave. A possible explanation for this is that following oxidation the pendant pyridyl group displaces a cyanide ligand and becomes coordinated to the metal centre. There is no other apparent reason why this alone of all the complexes should show irreversible behaviour on oxidation. The strong π-accepting effects of 2,2'-bipyrimidine are apparent, with the Ru$^{II/III}$ redox potential being at a more positive potential (+1.05 V vs. Ag/AgCl) than all the others (+0.70 to +0.89 V) because the metal centre is more electron-deficient. The expected ligand-centred reductions were not accessible within the limits of the potential window in H$_2$O.
A recent vibrational spectroscopic study by Kettle and co-workers of homoleptic tetra and hexacyanometallate complexes boldly attempts to present a framework for the comparison of transition metal carbonyl and cyanide ligands, on the basis of the isoelectronic nature of CO and CN$^-$. Relevant aspects of the study are discussed below. Kettle et al. highlight the significant electrostatic differences between the two ligands (CO is neutral whilst CN is anionic) and how this may facilitate cation binding, discussed in terms of “non-innocence” of the counter-ion. This can have significant effects on the electronic and vibrational spectra obtained; this effect is termed metallochromism and is investigated in §3.5.3. More subtle effects such as decreased dipole moment and increased quadrupole moment in metal cyanide complexes are investigated. These are postulated to give rise to unexpected band frequency ordering and intensity effects as first-order “ball-and-spring” models are proposed to be inaccurate for such complexes when quadrupolar interactions become significant (Figure 2.2.3.1).

A section of the paper is devoted to a DFT study exploring the dependence of vibrational band frequency on the formal charge of the metal centre in both metal carbonyl and cyanide complexes, an effect which is central to the sensitivity and utility of TRIR spectroscopy. The results suggest that the calculated vibrational band frequency varies non-linearly with the nuclear charge of the metal centre in both types of complex (Figure 2.2.3.2). As mentioned in Chapter 1 there is a plethora of TRIR data for metal carbonyl complexes which rely on this effect, but prior to this Thesis there have been very few TRIR studies of mononuclear transition metal cyanide complexes and a theoretical verification of the underlying concept of this Thesis is most timely and welcome!
Figure 2.2.3.1: Diagram of a possible metal-CN derived quadrupole coupling in a $M(CN)_3$ unit with $C_3v$ symmetry. Charge displacements are shown; violet corresponds to a charge decrease and blue corresponds to a charge increase.

Reproduced from Ref. 25

![Diagram of metal-CN coupling](image)

Figure 2.2.3.2: Calculated effect (DFT) of a variable positive charge placed 1.5 Å (black circles) or 2.0 Å (white circles) from the C atom on the stretching frequency of a) CN and b) CO in cm$^{-1}$. Reproduced from Ref. 25

![Graphs of CN and CO stretching frequencies](image)

[Ru(CN)$_4$(NN)]$^{2-}$ complexes have $C_{2v}$ symmetry and therefore these complexes should have four IR-active CN vibrations. In the solid state these are apparent as three closely spaced strong peaks between about 2030 and 2060 cm$^{-1}$, with a less intense peak at around 2090 cm$^{-1}$. In aqueous solution these are not all resolved; the three closely spaced bands merge to give a strong peak at 2050 – 2060 cm$^{-1}$, and a weaker one appearing as a shoulder about 20 – 30 cm$^{-1}$ higher (Table 2.2.3.1). Figures 2.2.3.1 and 4.3.3.1 show the ps-TRIR spectra of RuCNbpy and RuCNBPM in D$_2$O following 400 nm excitation respectively. In the spectrum of RuCNBPM, the parent
v(CN) vibrations are bleached, and weaker new bands appear at higher energy. Assuming that the most intense new transient corresponds to the most intense peak in the ground-state spectrum, there is a shift to higher energy of 52 cm\(^{-1}\) of this v(CN) vibration when the complex is in the \(^3\)MLCT excited state, as a consequence of loss of electron density on the Ru centre (which is transiently oxidised to Ru\(^{III}\)) and the corresponding decrease in Ru(d\(\pi\)) → CN(\(\pi^*\)) backbonding.

The low signal intensity of the transient v(CN) vibration compared to the ground state vibration is consistent with previous observations\(^9\) and suggests that the dipole of the C–N bonds is reduced in the \(^3\)MLCT excited state. This is reasonable as cyanide ligands formally have the negative charge on the C atom: transient oxidation of Ru\(^{II}\) to Ru\(^{III}\) will reduce the partial negative charge on the coordinated C atom and hence diminish the C–N dipoles, leading to weaker IR absorptions. The inherent weakness of the transient v(CN) signals is exacerbated by the fact that the transient signal at \(\sim2100\) cm\(^{-1}\) will be partially overlapping with the bleach of the weak, highest-energy v(CN) vibration in the ground state which will reduce the apparent intensity of the transient peak still further. Consequences of this are that only the most intense v(CN) transient peak can clearly be seen in typical TRIR spectra of tetracyanoruthenate (II) complexes and that these complexes can not be studied using two-colour IR ‘vibration-labelling’ experiments (§1.4.1.3).

Similar spectroscopic behaviour is shown by all of the complexes in D\(_2\)O, with a shift to higher energy of the most intense v(CN) vibration of between 46 and 52 cm\(^{-1}\) in the \(^3\)MLCT excited state, as evinced by a bleach of the ground-state vibrations and a less intense transient peak at higher energy. These values are the same within the limits of accuracy of the experiment, notwithstanding the 8-9 cm\(^{-1}\) resolution for the spectra and the low intensity of the transients. The fast timescale of this technique allowed us to determine accurately the lifetime of RuCNBP\(_M\), which could not be determined by luminescence methods as the emission was too weak and short-lived. §4.3.3 contains the TRIR difference spectrum for this complex, the kinetics of transient decays and
parent recovery and a detailed analysis of the results obtained. The rates derived from this data match closely and give $\tau = 3.4 \pm 0.3$ ns for the $^3$MLCT excited state in D$_2$O.

Two of the complexes were studied in CH$_3$CN to probe the effect of changing solvent on their transient IR behaviour. TRIR spectra of RuCNdmb and RuCNBPM measured in CH$_3$CN are similar to those recorded in D$_2$O, with the exception that the shift to higher energy of the main v(CN) vibration was smaller - being 36 cm$^{-1}$ for RuCNdmb (cf. 46 cm$^{-1}$ in D$_2$O) and 30 cm$^{-1}$ for RuCNBPM (cf. 48 cm$^{-1}$ in D$_2$O). This is probably a consequence of the fact that the most intense v(CN) vibration in the ground state is at higher energy in CH$_3$CN than in water (for RuCNdmb, 2068 vs. 2049 cm$^{-1}$; for RuCNBPM, 2068 vs. 2057 cm$^{-1}$), an effect which is ascribable to differences in hydrogen bonding with the solvent. The lifetime of the $^3$MLCT excited state of RuCNBPM in CH$_3$CN was found to be 250 (±50) ps, an order of magnitude shorter than the value observed in water, in keeping with the general trend expected for complexes of this series and the 'energy-gap law' (§1.3).

In order to simulate a higher-definition TRIR spectrum of the [Ru(CN)$_4$(NN)]$^{2-}$ unit, two spectra of RuCNbpy were taken in D$_2$O solution at $\Delta t = 100$ ps with data points approx 4-5 cm$^{-1}$ from each other and interpolated so as to simulate a spectrum with ca. 4 cm$^{-1}$ resolution, as shown in Figure 2.2.3.3. Even with an increased effective resolution only moderate insights may be taken from the spectrum, mainly in the clearer resolution of multiple features contributing to the transient peak profile between 2090 and 2130 cm$^{-1}$. This feature is best fitted to 2 components, centred at ca. 2099 and 2110 cm$^{-1}$ (fits not shown). Some spectra in later chapters (despite their lower resolution) clearly indicate the presence of two bands in similar positions. A particularly interesting example of this is the ps-TRIR spectrum of RuCN-BL$^1$-Rubpyamide in D$_2$O (Figure 4.5.3.1), which ostensibly shows the interconversion of 2 transient bands in this region.
<table>
<thead>
<tr>
<th>Complex</th>
<th>Solid State $v_{CN}$ (ground state)/ cm$^{-1}$</th>
<th>Solvent</th>
<th>TRIR Bleach centre/cm$^{-1}$</th>
<th>TRIR Transient centre/ cm$^{-1}$</th>
<th>$\Delta$/cm$^{-1}$</th>
</tr>
</thead>
<tbody>
<tr>
<td>RuCNbpy</td>
<td>2032, 2047, 2056, 2089</td>
<td>D$_2$O</td>
<td>2051</td>
<td>2103</td>
<td>52</td>
</tr>
<tr>
<td>RuCNdmb#</td>
<td>2034, 2040, 2054, 2090</td>
<td>D$_2$O, CH$_3$CN</td>
<td>2049</td>
<td>2095</td>
<td>46</td>
</tr>
<tr>
<td>RuCNphen</td>
<td>2032, 2038, 2047, 2088</td>
<td>D$_2$O</td>
<td>2051</td>
<td>2098</td>
<td>47</td>
</tr>
<tr>
<td>RuCNtpy</td>
<td>2041, 2047, 2091</td>
<td>D$_2$O</td>
<td>2051</td>
<td>2100</td>
<td>49</td>
</tr>
<tr>
<td>RuCNBPM#</td>
<td>2046, 2057, 2065, 2102</td>
<td>D$_2$O, CH$_3$CN</td>
<td>2057</td>
<td>2105</td>
<td>48</td>
</tr>
<tr>
<td>RuCNPBE</td>
<td>2028, 2044, 2085</td>
<td>D$_2$O</td>
<td>2048</td>
<td>2095</td>
<td>47</td>
</tr>
<tr>
<td>RuCNpypz</td>
<td>2028, 2045, 2052, 2092</td>
<td>D$_2$O</td>
<td>2050</td>
<td>2098</td>
<td>48</td>
</tr>
</tbody>
</table>

Table 2.2.3.1: TRIR data for the tetracyanoruthenate (II) complexes. *Solid-state spectra measured on powdered sample using diamond-ATR cell. #Solution FTIR and TRIR spectra of RuCNdmb and RuCNBPM are shown and discussed in detail in later Chapters (§3.2.3 and §4.3.3).

2.3. Discussion and Closing Points

This Chapter has presented highlights of the research literature on [Ru(CN)$_4$(bpy)]$^{2-}$ and related complexes in the context of photophysics, optical and vibrational spectroscopy and solvatochromism. The three principal advantages of [Ru(CN)$_4$(NN)]$^{2-}$ over [Ru(NN)$_3$]$^{2+}$ complexes have been discussed; the solvent-tunability of excited state energy, CN sites providing additional binding sites for metal ions and
Figure 2.2.3.3: Interpolated “high resolution” ps-TRIR spectrum of 
\([\text{Ru(CN)}_4(\text{bpy})]^{2-}\) in D$_2$O at $\Delta t = 100$ ps.

The sensitivity of $\nu$(CN) bands as spectroscopic reporters in the IR region of the electromagnetic spectrum.

We have extended the family of known tetracyanoruthenate (II) complexes with polypyridyl ligands and have documented their photophysical and electrochemical properties. All of the complexes reported in this Chapter have similar absorption and emission properties arising from the $^1$MLCT and $^3$MLCT transitions respectively, although synthetic manipulation of the complexes has altered their photophysical properties. Where studied the negative solvatochromism of the complexes is evident from the difference in absorption peak profile and emission energy, lifetime and quantum yield and confirms that solvatochromism is a potential external stimulus which may be used to modulate the excited state properties. TRIR spectroscopy has been successfully used to characterise the excited state properties of [Ru(CN)$_4$(NN)]$^{2-}$ complexes. These findings will be exploited and studied in more detail in Chapters 3 and 4 where [Ru(CN)$_4$(NN)]$^{2-}$ is used as a component of supramolecular assemblies designed to undergo intermetallic PEnT and PET between metal centres respectively.
and TRIR spectroscopy is used to directly monitor the photophysical processes \textit{in real time}.
2.4. References


Chapter 3:

Solvent Based Switching of the Gradient and Direction of Förster Photoinduced Energy Transfer in Polynuclear Complexes
Chapter 3: Solvent Based Switching of the Gradient and Direction of Förster Photoinduced Energy Transfer in Polynuclear Complexes

3.1. Introduction

In this Chapter a series of bimetallic complexes and their monometallic precursors are presented. They have been optimised for through-space (Förster) photoinduced energy transfer (PEnT) by the use of saturated organic bridging ligands, more precisely linkers between 2,2'-bipyridine molecules ligated to different metal centres. Here we develop supramolecular systems in which the direction and thermodynamic gradient of PEnT can be reversibly manipulated through changes in the solvent environment. The control of such processes by external stimuli is of considerable interest to the photophysical community and has applications in sensor devices and the development of molecular wires and photonics technologies.

Two relevant examples of bichromophoric solvent-switchable PEnT systems incorporating the [Ru(CN)₄]²⁻ moiety are discussed in the literature. Ward and co-workers studied the system {(NC)₄Ru(μ-BL)Ru(bpy)₂} (hereafter RU-CN-BL-Ru(bpy)) shown in Figure 3.1.1, consisting of a [Ru(CN)₄(bpy)]²⁻ unit covalently attached to a [Ru(bpy)₃]²⁺ unit through a flexible, 6-atom length polyoxoethylene chain. The lowest energy excited state of both chromophores may be formalised as Ru → bpyBL ³MLCT (where bpyBL is a 2,2'-bipyridine forming part of the bridging ligand BL). The [Ru(CN)₄(bpy)]²⁻ unit (hereafter RU-CN) exhibits negative solvatochromism (§1.2.4, Chapter 2) whereas the photophysics and spectroscopy of [Ru(bpy)₃]²⁺ (hereafter Ru(bpy)) are virtually solvent-independent.

The absorption spectra obtained for RU-CN-BL-Ru(bpy) in H₂O and (CH₃)₂SO, dimethylsulfoxide (hereafter DMSO) are shown in Figure 3.1.2. The main difference in the spectral profile in the two solvents is the prominence of a feature at λ_max = 550 nm in DMSO. This coincides with the lowest energy absorption of [Ru(CN)₄(bpy)]²⁻, also at λ_max = 550 nm (§2.2.1). In water the absorption of the RU-CN terminus is at
\( \lambda_{\text{max}} = 452 \text{ nm} \) and is obscured by the absorption from the Rubpy terminus, which shows little variation in the two solvents.

\[ \text{Figure 3.1.1: The bimetallic complex RuCN-BL-Rubpy.}^3 \]

The luminescence spectra of RuCN-BL-Rubpy measured in a series of DMSO:H\(_2\)O mixtures are shown in Figure 3.1.3. The emission maximum (following excitation of the Rubpy terminus) is invariant in the spectra with values in the range \( \lambda_{\text{max}} = 615 \) to 620 nm, indicating that emission from the Ru \( \rightarrow \) bpy \(^3\)MLCT excited state of the solvent-independent Rubpy terminus dominates the spectrum. The intensity and lifetime of the emission progressively decreases from H\(_2\)O to DMSO as the luminescence from the Rubpy unit becomes quenched more effectively by the RuCN terminus. As the solvent A. N. decreases the \([\text{Ru(CN)}_4(\text{bpy})]^2^-\) unit becomes both a
better energy acceptor (due to its lower energy $^3$MLCT state) and a better electron donor (due to its less positive oxidation potential, §1.4.1.2). The mechanism of quenching was proposed to be through-space PENt and a pictorial representation of the photophysical model is shown in Figure 3.1.4. Emission from the [Ru(CN)$_4$(bpy)]$^{2-}$ unit was found to be weak and undetected in either solvent. The authors suggested that this was obscured by the much more intense luminescence of the [Ru(bpy)$_3$]$^{2+}$ unit or efficient non-radiative decay pathways.

![Figure 3.1.3: The luminescence spectra of RuCN-BL-Rubpy in a series of DMSO:H$_2$O fractions (labelled). Reproduced from Ref. 3](image)

The authors proposed a photophysical model to rationalise the results observed. In H$_2$O the $^3$MLCT energies of the two termini are similar ($\Delta E = 230$ cm$^{-1}$ from luminescence spectroscopy of the monometallic model complexes) and the monoexponential emission lifetime of 360 ns is consistent with the lowest $^3$MLCT excited states on either terminus being in thermal equilibrium. In DMSO the negative solvatochromism of the RuCN unit is expected to result in a much lower $^3$MLCT energy than the Rubpy terminus ($\Delta E = 3600$ cm$^{-1}$ from luminescence spectroscopy of the monometallic model complexes). Emission is quenched fiftyfold in DMSO ($\phi_{\text{H}_2\text{O}} = 0.026$ whereas $\phi_{\text{DMSO}} = 0.0005$) and luminescence from the unquenched Rubpy termini is detected with a lifetime of 4 ns. The authors estimated the intermetallic distance from luminescence studies as 8 Å, representing a considerable degree of folding in the bridging ligand (the maximum distance possible being ca. 14 Å).
Indelli et al. linked the \([\text{Ru(CN)}_4(\text{bpy})]^{2-}\) unit with either one or two pyrene molecules \text{via} an ethane spacer through the bipyridyl ring at the 4,4' positions.\(^4\) The molecules \(\text{RuCNpyr}\) and \(\text{RuCN(pyr)}_2\) are shown in Figure 3.1.5. In contrast to the work of Simpson et al., the lowest energy excited states of the two chromophores in 5 and 6 are \(\text{Ru} \rightarrow \text{bpy}\) \(^3\)MLCT and pyrene \(^3\pi-\pi^*\) transitions. The MLCT transition energies of the \([\text{Ru(CN)}_4(\text{bpy})]^{2-}\) unit show the expected strong negative solvatochromism (§1.2.4) whereas the energies of the pyrene \(\pi-\pi^*\) transitions are virtually solvent-independent.

The absorption spectra of 5 in CH\(_3\)CN, CH\(_3\)OH and H\(_2\)O are shown in Figure 3.1.6. The \(^1\)MLCT absorption maximum varies from 392 nm in H\(_2\)O to 530 nm in CH\(_3\)CN,
exhibiting the expected negative solvatochromism of the RuCN unit (§1.2.1.4). Pyrene $^1\pi-\pi^*$ transitions occur below 350 nm and are not shown in the Figure below.

![Absorption spectra of RuCNpyr in H$_2$O (grey dotted line), CH$_3$OH (dashed line) and CH$_3$CN (solid line). Reproduced from Ref.4](image)

**Figure 3.1.6: Absorption spectra of RuCNpyr in H$_2$O (grey dotted line), CH$_3$OH (dashed line) and CH$_3$CN (solid line). Reproduced from Ref.4**

The emission spectra of RuCNpyr and RuCN(pyr)$_2$ were found to be strongly solvent-dependent. In H$_2$O their luminescence is extremely weak ($\lambda_{\text{max}} = 600$ nm), whereas in CH$_3$CN the emission ($\lambda_{\text{max}} = 780$ nm) is much stronger and is consistent with luminescence from a $^3$MLCT excited state. In CH$_3$OH the emission energy lies between the two extremes ($\lambda_{\text{max}} = 640$ nm) and the emission was found to decay with biexponential kinetics, with a major short-lived component and a minor longer-lived minor component. This is consistent with two states (one $^3$MLCT and one $^3\pi-\pi^*$) of similar energy both emitting.

In H$_2$O the $^3$MLCT excited state of the RuCN terminus lies higher in energy than the $^3\pi-\pi^*$ transition of the pyrene, and is completely quenched. Efficient energy transfer from the RuCN terminus to the pyrene group takes place. In CH$_3$CN the $^3$MLCT excited state energy of [Ru(CN)$_4$(bpy)]$^2^-$ is lower in energy due to the negative solvatochromism of transition metal cyanide complexes (§1.2.4) and lies lower in energy than the pyrene-based $^3\pi-\pi^*$ transitions. Emission is observed from the RuCN terminus.
In CH$_3$OH the $^3$MLCT excited state of the RuCN terminus and the $^3\pi-\pi^*$ excited state of the pyrene are of similar energy. The initially populated $^3$MLCT excited state undergoes energy transfer to the pyrene-based $^3\pi-\pi^*$ state but the process is reversible as the states are in equilibrium. The photophysics of the system is summarised in Figure 3.1.7, where the reversal of the ordering of the energies of the $^3$MLCT and $^3\pi-\pi^*$ excited states is illustrated.

Figure 3.1.7: The photophysics of RuCNpyr in H$_2$O (left), CH$_3$OH (centre) and CH$_3$CN (right). Reproduced from Ref.$^4$

These two studies illustrate the utility of the [Ru(CN)$_4$(bpy)]$^{2-}$ unit in systems where solvent composition can be used to tune the excited state ordering in either bimetallic or organic-inorganic systems. However there are limitations in the amount of useful information one can derive from luminescence spectra in these cases – especially in the case of RuCN-BL-Rubpy – as the broad, featureless nature of $^3$MLCT emission peaks and the low quantum yield and short lifetime of the [Ru(CN)$_4$(bpy)]$^{2-}$ unit, particularly in low A. N. solvents (§2.1.1) prevent conclusive understanding of the nature and mechanism of the PenT process. This is compounded by the fact that luminescence only provides information on the energy transfer process through the sensitised luminescence of the energy-accepting chromophore. Emission from this
process can be difficult to disentangle from other luminescence processes, such as
direct excitation of the acceptor chromophore, which for a strongly luminescent
chromophore such as [Ru(bpy)_3]^{2+} may dominate the spectrum.

For these reasons, in addition to luminescence we have employed ultrafast TRIR
spectroscopy to monitor the vibrational spectra of both termini of our bimetallic
complexes (based on the design of RuCN-BL-Rubpy) with the aim of directly
monitoring the Förster PEnT process in real time. §3.2, 3.3 and 3.4 detail the
photophysics and electrochemistry of the novel model complexes which the
bimetallics in §3.5, 3.6 and 3.7 are based upon. The results and implications of the
studies above are compared with those from this Thesis and analysed in more detail in
the Chapter Discussion (§3.8).

3.2. [Ru(CN)_4((5,5'-CH_3)_2bpy)]X_2 (X=K^+ or PPN^+), RuCNdmb

![Structure of RuCNdmb](image)

RuCNdmb is a member of the [Ru(CN)_4(NN)]^{2-} family which have been studied with
electrochemistry and absorption, luminescence and TRIR spectroscopy in Chapter 2.
Methyl groups have been introduced at the 5,5' positions on the bpy ligand in order to
more closely simulate the electronic environment of the metal centre in the dinuclear
complexes RuCN-BL^1-Rubpyamide (§3.5) and RuCN-BL^2-Rubpyamide (§3.6),
which are linked by an ethyl spacer. Spectroscopic studies have demonstrated that the
remote shielding of the cyanide ligand induces a downfield shift in the NMR spectrum
of the hydrogens in the 6,6' positions located above and below the ligand in
RuCNdmb. This is in contrast to the general aromatic ring current effect, which leads to an upfield shift in the $^1$H NMR signals for the other aromatic protons.\(^5\)

The similar ground and excited state properties of RuCNdmb and $[\text{Ru(CN)}_4(\text{bpy})]^2$ were compared in Chapter 2, and as expected RuCNdmb displays strong negative solvatochromism.

### 3.2.1. UV/visible Absorption and Luminescence Spectroscopy

The UV/visible absorption spectra of RuCNdmb have been measured in D$_2$O, H$_2$O, CH$_3$OH and CH$_3$CN (see Appendix §3.9 for spectra). The results are presented in Table 3.2.1.1. The observed negative solvatochromism of these complexes is due to SSDA interactions between the externally directed cyanide ligands and solvent molecules (§1.2.4). This phenomenon is responsible for a difference in the energy of the lowest energy $^1\text{MLCT}$ excited state of ca. 4600 cm$^{-1}$ between H$_2$O and CH$_3$CN.

<table>
<thead>
<tr>
<th>Solvent</th>
<th>$\lambda_{\text{max}}$ / nm</th>
<th>$\lambda_{\text{max}}$ (CH$_3$CN-H$_2$O) / nm</th>
<th>Assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>D$_2$O</td>
<td>293</td>
<td>293</td>
<td>IL $\pi$-$\pi^*$</td>
</tr>
<tr>
<td>H$_2$O</td>
<td>293</td>
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<td>$^1\text{MLCT}$</td>
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<tr>
<td>CH$_3$OH</td>
<td>293</td>
<td>293</td>
<td>$^1\text{MLCT}$</td>
</tr>
<tr>
<td>CH$_3$CN</td>
<td>297</td>
<td>297</td>
<td>$^1\text{MLCT}$</td>
</tr>
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<td>$\lambda_{\text{max}}$ (CH$_3$CN-H$_2$O)</td>
<td>4</td>
<td>41</td>
<td>$^1\text{MLCT}$</td>
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</table>

Table 3.2.1.1: Table of peak positions and assignments for UV/visible absorption spectra of RuCNdmb in D$_2$O, H$_2$O, CH$_3$OH and CH$_3$CN solutions.

The emission spectra of RuCNdmb were measured in D$_2$O, H$_2$O, CH$_3$OH and CH$_3$CN and the results are presented in Table 3.2.1.2. The broad, structureless emission observed is consistent with emission from $^3\text{MLCT}$ excited states, as observed in the family of $[\text{Ru(CN)}_4(\text{NN})]^2$ complexes in Chapter 2. The observed variation in $^3\text{MLCT}$ excited state energy ($\Delta E_{\text{H}_2\text{O}-\text{CH}_3\text{CN}} = 1800$ cm$^{-1}$) is significantly
smaller than the energy separation derived from the absorption spectrum ($\Delta E_{\text{H2O-CH3CN}} = 4600 \text{ cm}^{-1}$). In Ru $\rightarrow$ (NN) $^3$MLCT excited states partial oxidation of the metal centre leads to electron density being redistributed from CN$^-$ to Ru$^{\text{III}}$, decreasing the available electron density for donor-acceptor interactions (§1.2.4, also see §3.1.2).\(^6\)

<table>
<thead>
<tr>
<th>A. N.</th>
<th>$\lambda_{\text{max}}$ / nm</th>
<th>$\lambda_{\text{max}}$ Energy / cm$^{-1}$</th>
<th>$\tau$ / ns</th>
<th>$\phi$</th>
</tr>
</thead>
<tbody>
<tr>
<td>D$_2$O</td>
<td>54.8</td>
<td>614</td>
<td>16300</td>
<td>1020*</td>
</tr>
<tr>
<td>H$_2$O</td>
<td>54.8</td>
<td>614</td>
<td>16300</td>
<td>375 (+42)</td>
</tr>
<tr>
<td>CH$_3$OH</td>
<td>41.3</td>
<td>649</td>
<td>15400</td>
<td>210 (+15)</td>
</tr>
<tr>
<td>CH$_3$CN</td>
<td>19.3</td>
<td>729</td>
<td>14500</td>
<td>49 (+12)</td>
</tr>
</tbody>
</table>

Table 3.2.1.2: Table of peak positions, lifetimes and quantum yields from luminescence spectroscopy of RuCNdmb in D$_2$O, H$_2$O, CH$_3$OH and CH$_3$CN solutions. *The statistical error from curve fitting is $\pm 1$ ns, however this is smaller than a reasonable instrumental error.

The longer lifetime of RuCNdmb in D$_2$O relative to H$_2$O is attributed to a decrease in the non-radiative decay rate in deuterated solvents, as in the related complex [Ru(CN)$_4$(bpy)]$^{2-}$ (§2.1.2).\(^7\) The fluorescence lifetimes decrease as the Acceptor Numbers (A. N.) decrease and the emission energies decrease. This is in accordance with the ‘energy-gap law’, which states that the rate of non-radiative decay for a MLCT transition increases exponentially as the emission energy decreases (§1.3).\(^8\) The trend in quantum yield for the $^3$MLCT emission also follows this pattern, with $\phi_{\text{em}}$($\text{D}_2\text{O}$) > $\phi_{\text{em}}$($\text{H}_2\text{O}$) > $\phi_{\text{em}}$($\text{CH}_3\text{OH}$) $\approx$ $\phi_{\text{em}}$($\text{CH}_3\text{CN}$), in a similar manner to other [Ru(CN)$_4$(NN)]$^{2-}$ complexes (§2.1.2).

### 3.2.2. Cyclic Voltammetry

Cyclic voltammetry of RuCNdmb in CH$_3$CN was measured (see Appendix §3.9 for voltammograms) and the results are presented in Table 3.2.2.1 below. Features corresponding to one reduction and one oxidation couple are observed. Neither couple is easily discerned from the voltammogram. In particular the reduction was detected
as a discontinuity at a potential outside the solvent window (ca. -2 to +2 V vs. Ag/AgCl), so its reversibility was not determined. Upon application of potentials outside the solvent window the voltammogram shows features consistent with the products of solvent decomposition.

<table>
<thead>
<tr>
<th>$E_{1/2}/V$ (vs. Fc/Fc$^+$)</th>
<th>Reversibility</th>
<th>Assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>-0.03</td>
<td>Quasi-reversible</td>
<td>Ru$^{III}$</td>
</tr>
<tr>
<td>-2.57</td>
<td>Undetermined (vide supra)</td>
<td>(5,5'-CH$_3$)$_2$bpy$^{6/-}$</td>
</tr>
</tbody>
</table>

*Table 3.2.2.1: Table of potentials for redox couples and assignments for the cyclic voltammogram of RuCNdmb in CH$_3$CN solution.*

The oxidation potential is comparable to that of $[\text{Ru(bpy)}(\text{CN})_4]^{2-}$ in CH$_3$CN ($E_{1/2,\text{ox}} = -0.14$ V (vs. Fc/Fc$^+$)).$^9$ The reduction potential is similar to that of $[\text{Ru(bpy)}(\text{CN})_4]^{2-}$ in (CH$_3$)$_2$NCHO $E_{1/2,\text{red}} = -2.46$ V (vs. Fc/Fc$^+$).$^10$ It appears that the electron-donating effect of the methyl groups in the 5,5' positions makes the ligand reduction more difficult due to an increase in the energy of the ligand-based $\pi^*$ LUMO.

Attempts to study RuCNdmb in H$_2$O are hampered by the potential range permitted by the solvent (ca. -1 to +1 V vs. the Ag/AgCl). The bpy-based reduction is expected to have a similar potential in H$_2$O (ca. -2.5 V) and hence be outside the measurement range, whilst the oxidation would be expected to occur at a similar potential to that of $[\text{Ru(bpy)}(\text{CN})_4]^{2+}$. With $E_{1/2,\text{ox}} = +0.77$ V (vs. Fc/Fc$^+$),$^9$ the required potential vs. an Ag/AgCl reference electrode would be around +1.3 V.

### 3.2.3. Infrared Spectroelectrochemistry

Attempts to monitor the 1st reduction using FTIR spectroscopy on RuCNdmb in CH$_3$CN solution were unsuccessful as the process is chemically irreversible on the IR-OTTLE experimental timescale.
3.2.4. UV/visible/nIR Spectroelectrochemistry

Attempts to monitor the 1st reduction using UV/vis spectroscopy on RuCNdmb in CH₃CN solution were unsuccessful as the process is chemically irreversible on the UVVis-OTE experimental timescale.

3.2.5. FTIR and Time-Resolved Infrared Spectroscopy

The v(CN) region of the FTIR spectrum of RuCNdmb has been measured in CH₃CN and D₂O (see Appendix §3.9 for spectra). The complex has C₂ᵥ symmetry and should possess 4 IR active bands. It is clear from the spectral profile that several overlapping bands are present; visual inspection and least-squares fitting give at least 3 peak maxima in both CH₃CN and D₂O. These are at 2085, 2069, 2057 and 2044 cm⁻¹ in CH₃CN and at 2091, 2050 and 2030 cm⁻¹ in D₂O. These band positions are comparable with other [Ru(CN)₄(NN)]²⁻ complexes (Chapter 2). The difference in peak maxima is due to the hydrogen-bonding interaction of the cyanide ligands with solvent molecules and is further evidence for solvent-derived perturbation of vibronic energy levels in the complex.

Nanosecond and picosecond-TRIR spectra of RuCNdmb were measured in CH₃CN and D₂O solution. Figure 3.2.5.1 shows the ps-TRIR spectra obtained. Immediately after excitation the ground state bands are bleached and a transient feature appears at higher wavenumber. This feature has a maximum at 2108 cm⁻¹ in CH₃CN and 2096 cm⁻¹ in D₂O respectively. No evidence for vibrational cooling can be inferred from the ps-TRIR spectra. The bands formed do not decay over the picosecond timescale. The shifts of the v(CN) band to higher wavenumber are consistent with the formation of a Ru → (5,5'-CH₃₂bpy)³MLCT excited state.

As is observed in other members of the [Ru(CN)₄(NN)]²⁻ family (§2.2.3), the intensities of v(CN) bands in the IR region are much lower than comparable v(CO) bands and this can present spectral detection and signal-to-noise problems,
particularly with dilute or sparingly-soluble samples. Furthermore there is significant spectral overlap of the transient bands formed upon excitation with the high energy components of the ground-state \([\text{Ru(CN)}_4(\text{NN})]^2^-\) spectrum. These issues compound the difficulty in resolving single vibrational bands in both the ground and excited state spectral profiles and in some cases may limit the amount of information which can be extracted from the spectra owing to peaks not being clearly resolved (the PIRATE spectrometer has an instrumental resolution of ca. 8-9 cm\(^{-1}\), see Chapter 6 for more details).

The same spectral profile observed in the ps-TRIR spectra is reproduced on the nanosecond scale in both solvents. This profile decays with a time constant of 107 (±11) ns in CH\(_3\)CN. In D\(_2\)O the spectra decay with a biexponential time dependence, with time constants 38 (±9) and 430 (±130) ns (Figure 3.2.5.2). Lifetimes are calculated by averaging the values obtained from fitting both bleach and transient kinetic data. The TRIR results are summarised in Table 3.2.5.1.

<table>
<thead>
<tr>
<th></th>
<th>Ground State/ cm(^{-1})</th>
<th>Excited State/ cm(^{-1})</th>
<th>(\tau / \text{ns})</th>
</tr>
</thead>
<tbody>
<tr>
<td>D(_2)O</td>
<td>2091</td>
<td>2050</td>
<td>2030</td>
</tr>
<tr>
<td>CH(_3)CN</td>
<td>2089</td>
<td>2065</td>
<td>2057</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
</tbody>
</table>

*Table 3.2.5.1: Table of peak positions and lifetimes from TRIR spectra of RuCN\(_{dbm}\) in D\(_2\)O and CH\(_3\)CN solutions.*

3.2.6. RuCN\(_{dbm}\): Closing Points

The results from luminescence and TRIR spectroscopy confirm that the lowest-lying excited state of RuCN\(_{dbm}\) is Ru → \(((5,5'-(\text{CH}_3)_2\text{bpy})_3\text{MLCT}, with the excited state electron density located on the bipyridyl ring system. There is a strong negative solvatochromic effect on the absorption and emission energies of the MLCT transitions, and the "energy-gap law" may be invoked to rationalise the decrease in excited state lifetime and quantum yields as solvent A. N. decreases (§1.3).\(^8\) The range of \(^3\text{MLCT}\) energies exhibited by RuCN\(_{dbm}\) varies between 14500 cm\(^{-1}\)
(CH$_3$CN) and 16300 cm$^{-1}$ (D$_2$O) and this effect can be exploited to provide a means for modulating the excited state energy of the complex in polynuclear assemblies (vide supra). The emission quantum yield decreases by approximately an order of magnitude from H$_2$O to CH$_3$CN and hence the sensitising abilities of the complex are strongly solvent-dependent.

Our success in using $\nu$(CN) features in the infrared spectrum of [Ru(CN)$_4$(NN)]$^{2-}$ (§2.2.3 and §3.2.S) to probe excited state electronic structure unlocks the possibility of studying the excited state dynamics of RuCNdmmb (and supramolecular systems incorporating it) using TRIR on the picosecond and nanosecond timescales.

![TRIR spectra of RuCNdmmb in D$_2$O (left) and CH$_3$CN (right).](image)

**Figure 3.2.5.1:** TRIR spectra of RuCNdmmb in D$_2$O (left) and CH$_3$CN (right).

*Top:* Spectra recorded between 2 and 1000 ps after 400 nm laser excitation.

*Bottom:* Spectra recorded between 1 and 1000 ns after 355 nm laser excitation.
Figure 3.2.5.2: Kinetic traces from ns-TRIR spectra of RuCNdmb in D$_2$O (transient at 2097 cm$^{-1}$, top left and bleach at 2051 cm$^{-1}$, bottom left) and CH$_3$CN (transient at 2097 cm$^{-1}$, top right and bleach at 2104 cm$^{-1}$, bottom right). Fits shown in red are monoexponential.

3.3. [Ru(4,4'-CONEt$_2$)$_2$(bpy)$_2$Ru(1,2-bis-[5,5'-methyl-2,2'-bipyridyl]ethane)]Cl$_2$, Rubpyamide-BL$^1$

Rubpyamide-BL$^1$ is expected to display spectroscopic, electrochemical and photophysical characteristics similar to the parent complex [Ru(bpy)$_3$]$^{2+}$ and its amide-derivative analogue [Ru(4,4'-CO(NEt$_2$)bpy)$_2$(bpy)]$^{2+}$ (§1.4.1.1). The v(CO) bands of the amide groups provide a spectroscopic handle in the infrared region of the spectrum, probing the electron density on the bipyridyl rings in an analogous manner to studies performed by Meyer and co-workers (see §1.4.1.1 and references therein).
The replacement of the third bipyridine with the bridging ligand 1,2-bis-5-[5’-methyl-2,2’-bipyridyl]ethane adds a second bipyridyl ligand for metal complexation, facilitating the assembly of bimetallic complexes similar to RuCN-BL-Rubpy but with IR active spectroscopic reporter groups on both termini.

Unlike complexes based upon the [Ru(CN)₄]²⁻ moiety described in Chapter 2, Rubpyamide-BL¹ is expected to show little solvent-dependence on its electronic structure other than weak positive solvatochromism commonly observed for metal complexes (§1.2.2). As a result the emission energies of the complex are expected to be almost solvent-independent.

### 3.3.1. UV/visible Absorption and Luminescence Spectroscopy

UV/visible absorption spectra of Rubpyamide-BL¹ have been measured in CH₃CN and D₂O (see Appendix §3.9 for spectra). The spectral profile and peak positions are comparable to literature values for the parent complex [Ru(4,4’-(CONEt₂)₂bpy)₃]²⁺ and are presented in Table 3.3.1.1.¹² The peak wavelengths show much less solvent-dependence than those of RuCNdmb and other tetracyanoruthenate (II) complexes (Chapter 2 and §3.4.1).

<table>
<thead>
<tr>
<th>Solvent</th>
<th>λₘₐₓ / nm</th>
</tr>
</thead>
<tbody>
<tr>
<td>D₂O</td>
<td>&lt; 300</td>
</tr>
<tr>
<td></td>
<td>369</td>
</tr>
<tr>
<td></td>
<td>436</td>
</tr>
<tr>
<td></td>
<td>479</td>
</tr>
<tr>
<td>CH₃CN</td>
<td>&lt; 300</td>
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<tr>
<td></td>
<td>367</td>
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<tr>
<td></td>
<td>432</td>
</tr>
<tr>
<td></td>
<td>471</td>
</tr>
<tr>
<td>Assignment</td>
<td>IL π-π*</td>
</tr>
<tr>
<td></td>
<td>¹MC</td>
</tr>
<tr>
<td></td>
<td>¹MLCT</td>
</tr>
<tr>
<td></td>
<td>¹MLCT</td>
</tr>
</tbody>
</table>

**Table 3.3.1.1: Table of peak positions and assignments for UV/visible absorption spectra of Rubpyamide-BL¹ in D₂O and CH₃CN solutions.**

The emission profile of Rubpyamide-BL¹ has been measured in CH₃CN and D₂O (see Appendix §3.9 for spectra) and the results are presented in Table 3.3.1.2. The broad, structureless spectra are characteristic of emission from ³MLCT excited states and the spectra, lifetimes and quantum yields are comparable with those of [Ru(4,4’-
The slightly lower emission energy for Rubpyamide-BL in D$_2$O compared to CH$_3$CN (ca. 260 cm$^{-1}$) is due to increased stabilisation of the excited state by more polar solvent molecules.$^{13}$

<table>
<thead>
<tr>
<th>Solvent</th>
<th>$\lambda_{\text{max}}$ / nm</th>
<th>$\lambda_{\text{max}}$ Energy / cm$^{-1}$</th>
<th>$\tau$ / ns</th>
<th>$\varphi$</th>
</tr>
</thead>
<tbody>
<tr>
<td>D$_2$O</td>
<td>652</td>
<td>15340</td>
<td>367 (±41)</td>
<td>0.069</td>
</tr>
<tr>
<td>CH$_3$CN</td>
<td>641</td>
<td>15600</td>
<td>341 (±10)</td>
<td>0.061</td>
</tr>
</tbody>
</table>

Table 3.3.1.2: Table of peak positions, $^3$MLCT emission energies, lifetimes and quantum yields for luminescence spectra of Rubpyamide-BL in D$_2$O and CH$_3$CN solutions.

The $^3$MLCT excited state energy of Rubpyamide-BL (ca. 15500 cm$^{-1}$) lies within the range of $^3$MLCT excited state energies determined for RuCNdmb (16300 - 14500 cm$^{-1}$, §3.2.1). It follows that the location of the lowest energy excited state in a dinuclear complex containing the RuCNdmb and Rubpyamide-BL moieties should be tuneable by changing the solvent composition, provided that combination of the two fragments does not significantly perturb their electronic structure.

3.3.2. Cyclic Voltammetry

Rubpyamide-BL has been studied using cyclic voltammetry in CH$_3$CN (see Appendix §3.9 for the voltammogram) and the results are presented in Table 3.3.2.1. There is one reversible oxidation which arises from the Ru$^{II/III}$ couple, a large irreversible feature at positive potential (due to oxidation of the amide groups) and three reversible reduction couples arising from the bpy-based ligands. The potentials for the oxidation and the 1$^{\text{st}}$ and 2$^{\text{nd}}$ reductions are similar to those reported for [Ru(4,4'-($\text{CONEt}_2$)$_2$bpy)$_3$]$^{2+}$: $E_{\text{V,ox}} = +0.93$ V, $E_{\text{V,red1}} = -1.54$ V and $E_{\text{V,red2}} = -1.68$ V (all vs. Fc/Fc', §1.4.1.1).$^{14}$
Table 3.3.2.1: Table of potentials for redox couples and assignments for the cyclic voltammogram of Rubpyamide-BL\textsuperscript{1} in CH\textsubscript{3}CN solution.

<table>
<thead>
<tr>
<th>(E_{\text{red}}/V) (vs. Fe/Fe\textsuperscript{+})</th>
<th>Reversibility</th>
<th>Assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>+0.95</td>
<td>Reversible</td>
<td>Ru\textsuperscript{II/III}</td>
</tr>
<tr>
<td>((E_{p})) +0.78</td>
<td>Irreversible</td>
<td>((4,4'-(\text{CONEt}_2)_2\text{bpy})^{0/+})</td>
</tr>
<tr>
<td>-1.57</td>
<td>Reversible</td>
<td>((4,4'-(\text{CONEt}_2)_2\text{bpy})^{0/-})</td>
</tr>
<tr>
<td>-1.76</td>
<td>Reversible</td>
<td>((4,4'-(\text{CONEt}_2)_2\text{bpy})^{0/-})</td>
</tr>
<tr>
<td>-2.15</td>
<td>Reversible</td>
<td>((\text{BL}_1)^{0/+})</td>
</tr>
</tbody>
</table>

The 1\textsuperscript{st} and 2\textsuperscript{nd} reductions occur closely together (\(\Delta E_{\text{red2-red1}} = -0.19\) V), indicating that they take place on the amide-substituted bpy ligands, which are expected to have lower-energy \(\pi^*\) orbitals than BL\textsuperscript{1} due to the electron-withdrawing nature of the amide groups. As there are two identical \((4,4'-(\text{CONEt}_2)_2\text{bpy})\) ligands in the complex, the difference between the 1\textsuperscript{st} and 2\textsuperscript{nd} reduction potentials corresponds to the increased repulsion from adding another electron to an already reduced complex.\textsuperscript{14} The 3\textsuperscript{rd} reduction takes place at a much more negative potential (\(\Delta E_{\text{red3-red2}} = -0.39\) V) and is centred on the ligated bipyridyl of BL\textsuperscript{1}. A 4\textsuperscript{th} reduction centred on the bridging bipyridyl would be expected at more negative potentials, but was not observed in the potential window of CH\textsubscript{3}CN (ca. +2 to -2 V).

### 3.3.3. Infrared Spectroelectrochemistry

IR spectroelectrochemistry was used to monitor the change in spectral profile of the amide v(CO) bands of Rubpyamide-BL\textsuperscript{1} in CH\textsubscript{3}CN as each of the redox couples in Table 3.3.2.1 was stepped across.

**Ru\textsuperscript{II/III} Oxidation**

The change in spectral profile (not shown) is irreversible and the only spectroscopic change observed is a decrease in the absorbance of the parent band. This is attributed
to the decomposition of the amide groups at less positive potentials than that required
to oxidise the metal centre.

\((4,4'-(\text{CONEt}_2)_2\text{bpy})^{0/\text{-}}\) Reduction Couples

Figure 3.3.3.1 shows the spectra obtained as the first and second reductions occur. Between -0.7 and -1.2 V (vs. Ag) the parent feature at 1639 cm\(^{-1}\) decreases in absorbance, its profile broadens and its centre shifts to lower wavenumber. Considerable overlap with a solvent band makes it difficult to extract further information from the raw spectra. This is consistent with an increase in electron density on the \((4,4'-(\text{CONEt}_2)_2\text{bpy})\) ligands and confirms that the first two reductions are taking place there, as predicted from the cyclic voltammogram (\textit{vide supra}). The spectra in Figure 3.3.3.1 are broad and structureless, however the subtraction spectra presented in Figure 3.3.3.2 allow the changes in spectral profile to be visually inspected as the reductions take place. Two peaks appear in the spectra at \textit{ca.} 1622 and 1582 cm\(^{-1}\). The 1622 cm\(^{-1}\) band grows in only during the 1\textsuperscript{st} reduction (Figure 3.3.3.2, top) and the 1582 cm\(^{-1}\) band grows in only during the 2\textsuperscript{nd} reduction (Figure 3.3.3.2, bottom).

This suggests that each reduction is \textit{localised} on the IR timescale (§1.3) and takes place on a particular ligand, so that features arising from neutral and reduced \((4,4'-(\text{CONEt}_2)_2\text{bpy})\) are present throughout the experiment. If the electron were \textit{delocalised} over both \((4,4'-(\text{CONEt}_2)_2\text{bpy})\) ligands then only one peak would appear as the reduction occurred. Previous photophysical and electrochemical studies on \textit{bis}-heteroleptic Ru(II) polypyridine complexes have observed electron localisation on single bipyridyl ligands.\textsuperscript{14,15}

These data are consistent with the findings from the CV measurements that the individual reductions themselves are not isolated processes and there may be singly and doubly reduced species present in the spectra ascribed to the ‘first reduction’. Coulometric studies on \([\text{Ru}((4,4'-(\text{CONEt}_2)_2\text{bpy})_3]^2+\) show that the first reduction involves the addition of 0.8 electrons to the complex and the second adds 1.4.\textsuperscript{16}
Figure 3.3.3.1: FTIR spectra of the 1st and 2nd reductions of Rubpyamide-BL in CH₃CN. The initial spectrum (A, red), the spectrum obtained after the 1st reduction (B, cyan) and the final spectrum (C, green) are highlighted. Changes in spectral profile are reversible.

Figure 3.3.3.2: Difference IR spectra of the 1st and 2nd reductions of Rubpyamide-BL in CH₃CN. Top: Difference spectrum obtained from Figure 3.3.3.1 by subtracting the spectrum B from spectrum A. Middle: spectrum C – spectrum A. Bottom: spectrum C – spectrum B.
The changes in spectral profile as the complex undergoes its 3rd reduction are irreversible under the IR-OTTLE experimental conditions.

3.3.4. UV/visible/nIR Spectroelectrochemistry

**RuII/III Oxidation**

The change in the electronic absorption profile of **Rubpyamide-BL** was monitored in CH$_3$CN as the complex was oxidised. The change (not shown) was found to be irreversible and is presumably due to oxidative decomposition of the amide groups (vide supra).

**(4,4'-(CONEt$_2$)$_2$bpy)$^{0+/+}$ Reduction Couples**

Figure 3.3.4.1 shows UV/vis/nIR spectra for **Rubpyamide-BL** in CH$_3$CN as the complex undergoes its first and second reductions at -1.00 and -1.25 V vs. Ag/AgCl respectively. The first reduction is accompanied by small spectral changes; shoulders at 368 and 537 nm grow in. Major spectral changes occur during the 2nd reduction. Peaks at 347, 511 and 546 nm and a very broad feature with a maximum at 1308 nm appear. These are tentatively assigned to a $\pi-\pi^*$ transition, two $^1$MLCT transitions and various electronic transitions of the reduced $(4,4'-(CONEt$_2$)$_2$bpy)$ ligand respectively. The features corresponding to $^1$MLCT transitions in the original complex at 432 and 466 nm both decrease. These changes are consistent with an increase in electron density on the $(4,4'-(CONEt$_2$)$_2$bpy)$ ligands. The changes in spectral profile are reversible. Analogous behaviour has been reported for the 1st and 2nd reductions of the complex [Ru$(4,4'-(CONEt$_2$)$_2$bpy)$_3$]$^{2+}$.14
3.3.4.1: The UV/visible/nIR spectra of the reduction of Rubpyamide-BL in CH$_3$CN. Top: The 1st reduction at -1.00 V; inset: magnification of the region containing major changes. Bottom: The 2nd reduction at -1.25 V; inset: magnification of the region containing major changes. The initial spectrum (thick black) is emphasised in each case. Potentials are quoted vs. Ag/AgCl.

3.3.5. FTIR and Time-Resolved Infrared Spectroscopy

The amide v(CO) region of the FTIR spectrum of Rubpyamide-BL has been measured in CH$_3$CN and D$_2$O (see Appendix §3.9 for spectra). There is a single peak present at 1638 cm$^{-1}$ in CH$_3$CN and at 1612 cm$^{-1}$ in D$_2$O. The difference in the peak
centres is due to the hydrogen-bonding interaction of the amide groups with solvent molecules.

TRIR spectra of Rubpyamide-BL were measured in CH$_3$CN and D$_2$O solution. Figure 3.3.5.2 shows the ps-TRIR spectra obtained. Immediately after excitation the ground state bands bleach and transient bands at lower wavenumber appear at 1621 cm$^{-1}$ in CH$_3$CN and 1591 cm$^{-1}$ in D$_2$O respectively. In contrast with similar ps-TRIR studies on transition metal complexes with ester-substituted bpy ligands, there is no spectral evidence of vibrational cooling processes. This suggests that vibrational cooling occurs within 2 ps of the laser flash, or that the heavy overlap between bleach and transient mask its observation in the ps-TRIR spectra and kinetics shown. The bands formed do not decay over the picosecond timescale. The peak shifts to lower wavenumber are consistent with the formation of a $^3$MLCT excited state which is delocalised over the two (4,4'-((CONEt$_2$)$_2$)bpy) ligands on the IR timescale ($§$1.3).

The same spectral profile is observed on the nanosecond scale in both CH$_3$CN and D$_2$O, which decays with time constants of 316 (±17) ns in CH$_3$CN and 515 (±42) ns in D$_2$O respectively, averaged over bleach and transient features (Figure 3.3.5.3). The results are summarised in Table 3.3.5.1.

<table>
<thead>
<tr>
<th>Solvent</th>
<th>Bleach / cm$^{-1}$</th>
<th>Transient / cm$^{-1}$</th>
<th>$\tau$ / ns</th>
</tr>
</thead>
<tbody>
<tr>
<td>D$_2$O</td>
<td>1612</td>
<td>1591</td>
<td>515 (±42)</td>
</tr>
<tr>
<td>CH$_3$CN</td>
<td>1638</td>
<td>1621</td>
<td>316 (±17)</td>
</tr>
</tbody>
</table>

Table 3.3.5.1: Table of peak positions and lifetimes for ns-TRIR spectra of Rubpyamide-BL$^1$ in D$_2$O and CH$_3$CN solutions.
Figure 3.3.5.3: TRIR spectra of Rubpyamide-BL¹ in D₂O (left) and CH₃CN (right).

Top: ps-TRIR spectra recorded between 2 and 1000 ps after 400 nm excitation.
Bottom: ns-TRIR spectra recorded between 1 and 1000 ns after excitation.

Figure 3.3.5.4: Kinetic traces from TRIR spectra of Rubpyamide-BL¹ in D₂O (transient at 1593 cm⁻¹, top left and bleach at 1617 cm⁻¹, bottom left) and CH₃CN (transient at 1616 cm⁻¹, top right and bleach at 1636 cm⁻¹, bottom right). The fits shown are single exponentials.
3.3.6. Rubpyamide-BL\(^1\): Closing Points

Electrochemistry confirms that Rubpyamide-BL\(^1\) has a similar electronic structure to the complex \([\text{Ru}(4,4'\text{-}(\text{CONEt}_2)_2\text{bpy})_3]^2^+\), with the important difference that the bridging ligand's reduction couple is less facile. IR Spectroelectrochemistry of the complex in CH\(_3\)CN suggests that the first and second reduction take place sequentially and reduce the complex in a *localised* manner.

In comparison, the results from luminescence and TRIR spectroscopy suggest that the lowest-lying excited state of Rubpyamide-BL\(^1\) corresponds to \(\text{Ru} \rightarrow (4,4'\text{-}(\text{CONEt}_2)_2\text{bpy})^3\text{MLCT}\), with the presence of only one transient in the TRIR spectra strongly suggesting that the electron is *delocalised* over both amide-substituted ligands. As expected for transition metal complexes, there is a small positive solvatochromic effect on the absorption and emission energies of the MLCT transitions (§1.2.2). These apparently conflicting reports on the localisation of MLCT excited states are ostensibly due to the *timescale* of the exchange of excited state charge density distributions being faster or slower than the IR timescale (§1.3).

Rubpyamide-BL\(^1\) has excited state energies (15300 to 15600 cm\(^{-1}\)) which fall within the range displayed by RuCNdmb in the solvents studied (14500 to 16300 cm\(^{-1}\), Chapter 2 and §3.2.1). Thus the direction and gradient of photoinduced energy transfer should be tunable by altering the solvent composition. The emission quantum yield and \(^3\text{MLCT}\) excited state lifetimes show a weak positive solvatochromism which is expected for these complexes.
Rubpyamide-BL$^2$ is derived from Rubpyamide-BL$^1$ with a few important differences. The link between the polypyridyls of the bridging ligand has been modified from a conformationally flexible ethyl chain to a closed ring resembling a macrocycle, and the bpy rings are substituted in the 4,4' positions (as opposed to the 5,5' substitution in Rubpyamide-BL$^1$). This places further restrictions on the ability of the ligand to adopt different spatial geometries and subsequently constricts the free movement of Rubpyamide-BL$^2$. The change in substitution pattern on BL$^2$ is not expected to make a significant difference to the photophysics or electrochemistry of the complex, as the (4,4'-CONEt$_2$hbpy) ligands possess lower energy MLCT excited states and less negative reduction potentials than the bridging ligand.

Uv/vis absorption and emission spectroscopy, electrochemistry and TRIR spectroscopy have been utilised to study the electronic structure of Rubpyamide-BL$^2$ and to compare it with its conformationally unrestricted analogue (Rubpyamide-BL$^1$). As with Rubpyamide-BL$^1$, little or no solvent-dependence on its electronic structure is expected.
3.4.1. UV/visible Absorption and Luminescence Spectroscopy

UV/visible absorption spectra of Rubpyamide-Bl\(^2\) have been measured in CH\(_3\)CN and D\(_2\)O (see Appendix §3.9 for spectra). The spectral profile and peak positions are similar to Rubpyamide-Bl\(^1\) (§3.3.1) and literature values for the complex [Ru(4,4'-(CONEt\(_2\))\(_2\)bpY)]\(^{2+}\) (§1.4.1.1) and are presented in Table 3.4.1.1.\(^{12}\)

<table>
<thead>
<tr>
<th>Solvent</th>
<th>(\lambda_{\text{max}} / \text{nm})</th>
</tr>
</thead>
<tbody>
<tr>
<td>D(_2)O</td>
<td>&lt;300 \hspace{10mm} 360 \hspace{10mm} 437 \hspace{10mm} 471</td>
</tr>
<tr>
<td>CH(_3)CN</td>
<td>&lt;300 \hspace{10mm} 357 \hspace{10mm} 437 \hspace{10mm} 468</td>
</tr>
<tr>
<td>Assignment</td>
<td>IL (\pi-\pi^*) \hspace{10mm} (1^\text{MC}) \hspace{10mm} (1^\text{MLCT}) \hspace{10mm} (1^\text{MLCT})</td>
</tr>
</tbody>
</table>

*Table 3.4.1.1: Table of peak positions and assignments for UV/visible absorption spectra of Rubpyamide-Bl\(^2\) in D\(_2\)O and CH\(_3\)CN solutions.*

Emission spectra of Rubpyamide-Bl\(^2\) have been measured in CH\(_3\)CN and D\(_2\)O (see Appendix §3.9 for spectra). A broad, featureless spectrum characteristic of \(3^\text{MLCT}\) emission is obtained, with lifetimes and quantum yields similar to Rubpyamide-Bl\(^1\) (§3.3.1) and [Ru(4,4'-CO\(_2\)Et\(_2\)bpY)]\(^{2+}\) (§1.4.1.1).\(^{12}\) The results are presented in Table 3.4.1.2. Rubpyamide-Bl\(^2\) also displays the expected slight positive solvatochromism in D\(_2\)O compared to CH\(_3\)CN.\(^{13}\)

<table>
<thead>
<tr>
<th>Solvent</th>
<th>(\lambda_{\text{max}} / \text{nm})</th>
<th>(\lambda_{\text{max}} \text{ Energy / cm}^{-1})</th>
<th>(\tau / \text{ns})</th>
<th>(\phi)</th>
</tr>
</thead>
<tbody>
<tr>
<td>D(_2)O</td>
<td>654</td>
<td>15290</td>
<td>457 (±14)</td>
<td>0.054</td>
</tr>
<tr>
<td>CH(_3)CN</td>
<td>647</td>
<td>15460</td>
<td>462 (±8)</td>
<td>0.054</td>
</tr>
</tbody>
</table>

*Table 3.4.1.2: Table of peak positions, \(3^\text{MLCT}\) emission energies, lifetimes and quantum yields for luminescence spectra of Rubpyamide-Bl\(^2\) in D\(_2\)O and CH\(_3\)CN solutions.*

3-26
Rubpyamide-BL\textsuperscript{2} has been studied using cyclic voltammetry in CH\textsubscript{3}CN (see Appendix §3.9 for the voltammogram) and the results are presented in Table 3.3.2.1. As with Rubpyamide-BL\textsuperscript{1}, there is one reversible oxidation, a large irreversible feature at positive potential (due to oxidation of the amide groups) and three reversible reduction couples arising from the bpy-based ligands. The potentials for the oxidation and the 1\textsuperscript{st} and 2\textsuperscript{nd} reductions are similar to those reported for Rubpyamide-BL\textsuperscript{1} (§3.3.2) and [Ru(4,4'-(CONEt\textsubscript{2})\textsubscript{2}bpy)\textsubscript{3}]\textsuperscript{2+} (§1.4.1.1).\textsuperscript{14}

<table>
<thead>
<tr>
<th>$E_w$/V (vs. Fe/Fe\textsuperscript{+})</th>
<th>Reversibility</th>
<th>Assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>+0.95</td>
<td>Reversible</td>
<td>Ru\textsuperscript{II/III}</td>
</tr>
<tr>
<td>(Ep) +0.78</td>
<td>Irreversible</td>
<td>(4,4'-(CONEt\textsubscript{2})\textsubscript{2}bpy)\textsuperscript{0/+}</td>
</tr>
<tr>
<td>-1.57</td>
<td>Reversible</td>
<td>(4,4'-(CONEt\textsubscript{2})\textsubscript{2}bpy)\textsuperscript{0/-}</td>
</tr>
<tr>
<td>-1.76</td>
<td>Reversible</td>
<td>(4,4'-(CONEt\textsubscript{2})\textsubscript{2}bpy)\textsuperscript{0/-}</td>
</tr>
<tr>
<td>&lt;-2.00</td>
<td>Reversible</td>
<td>(BL\textsuperscript{1})\textsuperscript{0/-}</td>
</tr>
</tbody>
</table>

*Table 3.4.2.1: Table of potentials for redox couples and assignments for the cyclic voltammogram of Rubpyamide-BL\textsuperscript{2} in CH\textsubscript{3}CN solution.*

The 1\textsuperscript{st} and 2\textsuperscript{nd} reductions occur closely to each other ($\Delta E_{red2-red1} = -0.19$ V), indicating that they take place on the amide-substituted bpy ligands, which are expected to have lower-energy $\pi^*$ orbitals than BL\textsuperscript{1} due to the electron-withdrawing nature of the amide groups. As there are two identical (4,4'-(CONEt\textsubscript{2})\textsubscript{2}bpy) ligands in the complex, the difference between the 1\textsuperscript{st} and 2\textsuperscript{nd} reduction potentials corresponds to the increased repulsion from adding another electron to an already reduced complex.\textsuperscript{14} As with Rubpyamide-BL\textsuperscript{1}, the 3\textsuperscript{rd} reduction takes place at a much more negative potential and is centred on the ligated bipyridyl of BL\textsuperscript{2}.  

3-27
3.4.3. FTIR and Time-Resolved Infrared Spectroscopy

The FTIR spectrum of Rubpyamide-BL\textsuperscript{2} has been measured in CH\textsubscript{3}CN and D\textsubscript{2}O in the amide ν(CO) region (see Appendix §3.9 for spectra). As for Rubpyamide-BL\textsuperscript{1}, there is a single peak present at 1636 cm\textsuperscript{-1} in CH\textsubscript{3}CN and at 1614 cm\textsuperscript{-1} in D\textsubscript{2}O. The difference in the peak centre positions is due to the hydrogen-bonding interaction of the amide groups with solvent molecules.

Nanosecond and picosecond TRIR spectroscopy were performed on Rubpyamide-BL\textsuperscript{2} in CH\textsubscript{3}CN and D\textsubscript{2}O solutions. Figure 3.4.3.1 shows the ps-TRIR spectra obtained. The spectra collected exhibit significant noise levels due to instability of the probe laser during the experimental time allocation at RAL during which the data were collected. Upon excitation the ground state bands bleach and transient bands to lower wavenumber appear at 1617 cm\textsuperscript{-1} in CH\textsubscript{3}CN and 1596 cm\textsuperscript{-1} in D\textsubscript{2}O respectively. No spectral evidence for vibrational cooling following laser excitation is observed in the spectra. The bands formed do not appear to decay over the picosecond timescale. As with Rubpyamide-BL\textsuperscript{1}, the shifts of the ν(CO) band to lower wavenumber are consistent with charge transfer from the ruthenium atom to the two (4,4'-CONEt\textsubscript{2})zbpy) ligands, forming a delocalised \textsuperscript{3}MLCT excited state (as probed on the IR timescale, §1.3).

A similar spectral profile is observed on the nanosecond scale for each solvent. The peaks decay with biexponential profiles, with time constants (averaged between bleach and transient data) of 7.8 (±1.4) and 500 (±48) ns in CH\textsubscript{3}CN. In D\textsubscript{2}O a biexponential decay is also observed with time constants 7.5 (±1.1) and 470 (±51) ns (Figure 3.4.3.2). The results obtained are summarised in Table 3.4.3.1.
Table 3.4.3.1: Table of peak positions and lifetimes for TRIR spectra of Rubpyamide-BL² in D₂O and CH₃CN solutions.

<table>
<thead>
<tr>
<th>Solvent</th>
<th>Bleach / cm⁻¹</th>
<th>Transient / cm⁻¹</th>
<th>τ / ns</th>
</tr>
</thead>
<tbody>
<tr>
<td>D₂O</td>
<td>1614</td>
<td>1596</td>
<td>7.5 (±1.1), 470 (±51)</td>
</tr>
<tr>
<td>CH₃CN</td>
<td>1636</td>
<td>1617</td>
<td>7.8 (±1.4), 500 (±48)</td>
</tr>
</tbody>
</table>

Figure 3.4.3.1: TRIR spectra of Rubpyamide-BL² in D₂O (left) and CH₃CN (right).

Top: ps-TRIR spectra recorded between 2 and 1000 ps after 400 nm excitation.

Bottom: ns-TRIR spectra recorded between 1 and 300 ns after 355 nm excitation.

The spectra recorded were of poor quality due to instability of the PIRATE apparatus’ probe laser during the experimental period in which the data were collected.
Figure 3.4.3.2: Kinetic traces from ns-TRIR spectra of Rubpyamide-BL$_2$ in D$_2$O (transient at 1594 cm$^{-1}$, top left and bleach at 1617 cm$^{-1}$, bottom left) and CH$_3$CN (transient at 1624 cm$^{-1}$, top right and bleach at 1643 cm$^{-1}$, bottom right). The traces in D$_2$O are fitted to dual exponentials whilst the traces in CH$_3$CN are fitted to a single exponential.

3.4.4. Rubpyamide-BL$_2$: Closing Points

The results from luminescence and TRIR spectroscopy suggest that the lowest-lying excited state of Rubpyamide-BL$_2$ is similar to Rubpyamide-BL$_1$ and corresponds to Ru $\rightarrow$ (4,4'-(CONEt$_2$)$_2$bpy) $^3$MLCT. Rubpyamide-BL$_2$ has excited state energies between 15300 and 15500 cm$^{-1}$, which fall within the range displayed by RuCN$^{mb}$ in the solvents studied (14500 to 16300 cm$^{-1}$, Chapter 2 and §3.2.1). Thus the direction and gradient of photoinduced energy transfer in dyads containing these two chromophores should be controllable through alteration of the solvent environment.

An interesting contrast between the two Rubpyamide-BL$_x$ compounds is the biexponential decay of the ns-TRIR spectra in Rubpyamide-BL$_2$. The longer-lived
component (ca. 480 ns) corresponds fairly well to the luminescence lifetime (ca. 460 ns), but there is no obvious photophysical process which explains the shorter-lived component (ca. 8 ns). It may be a consequence of the synthesis of the BL\textsuperscript{2} ligand giving rise to two different orientations of the bridging ligand. This possibility is explored further in the conformational analysis of \textbf{RuCN-BL\textsuperscript{2}-Rubpyamide} (§3.6.5). This additional component may be present due to an impurity as a result of the preparation of BL\textsuperscript{2} which also produces a number of polypyridyl by-products.\textsuperscript{19} Alternatively the high analyte concentrations employed in order to obtain good signal-to-noise ratios may have induced self-quenching or aggregation (see Chapter 6 for more information on the TRIR experimental setup). The noise observed in the spectra at longer time delays could be due to some acoustic process producing shockwaves in the sample cell which distort the flow of the analyte solution and introduce artefacts into the spectrum. Attempts to collect better quality spectra during other allocations of experimental time using the PIRATE apparatus were unsuccessful.

Having studied the photophysical and electrochemical properties of the model complexes \textbf{RuCNdmb}, \textbf{Rubpyamide-BL\textsuperscript{1}} and \textbf{Rubpyamide-BL\textsuperscript{2}} this study will now examine the bimetallic complexes prepared by complexing the fragments [Ru(CN)\textsubscript{4}]\textsuperscript{2-}, [Ru(4,4’-(CONE\textsubscript{2}h2bpy)\textsubscript{2}]\textsuperscript{2+} and [ReCl(CO)\textsubscript{3}] to the bridging ligands BL\textsuperscript{1} and BL\textsuperscript{2}. These complexes should be regarded as second-generation analogues of the \textbf{RuCN-BL-Rubpy} system described earlier (§3.1).
RuCN-BL\textsuperscript{1}-Rubpyamide may be considered a fusion of the two mononuclear complexes RuCN\textsubscript{dmb} (§3.2) and Rubpyamide-BL\textsuperscript{1} (§3.3) and has been designed to exploit properties of these two building blocks. Both exhibit lowest-lying 3MLCT excited states as characterised by absorption, luminescence and TRIR spectroscopy. SSDA interactions modulate the 3MLCT energy of RuCN\textsubscript{dmb} so that it is ca. 700 cm\textsuperscript{-1} higher than the 3MLCT energy of Rubpyamide in high A. N. solvents such as D\textsubscript{2}O, but ca. 190 cm\textsuperscript{-1} lower in solvents such as CH\textsubscript{3}CN (\textit{vide supra}). As a result the direction and gradient for PEnT in this system should be reversibly controllable through the tuning of solvent A. N.

There are some key differences between this system and the previously studied RuCN-BL-Rubpy (§4.1). The modifications to append amide groups to the 4,4' positions of the bipyridyl ligands of RuCN-BL\textsuperscript{1}-Rubpyamide lowers the energy of Ru \textrightarrow (4,4'-(CONEt\textsubscript{2})\textsubscript{2}bpy) 3MLCT transitions so that they are the lowest-lying excited states in this system. By contrast the lowest energy transition resides on the bridging ligand for both chromophores RuCN-BL-Rubpy prepared by Simpson \textit{et al.}\textsuperscript{3} Secondly the change of bridging ligand length (shortened from 6 in RuCN-BL-Rubpy to 4 atoms here) and composition (from a POE chain to an ethyl linker) may affect the interchromophoric distance and conformational distribution sampled during
these solution experiments. These factors may have an effect on the rate and efficiency of through-space PEnT in this system.

The BL\textsuperscript{1} bridging ligand 1,2-bis-5-[5'-methyl-2,2'-bipyridyl]ethane allows the intermetallic distance to be modulated. Efficient PEnT via the Förster mechanism requires that the centres are spatially proximal and that there is good orientational alignment between the dipoles of the two chromophores.\textsuperscript{20} The Dexter mechanism is unlikely to operate in this system because the saturated CH\textsubscript{2}CH\textsubscript{2} linker prevents direct electronic communication between the two metal centres.

The v(CO) bands of the amide groups and the v(CN) bands of the cyanide ligands provide complementary vibrational handles in different regions of the IR spectrum, enabling real time direct monitoring of the evolution of the excited state structure on both metal centres. The insights into excited state electronic structure obtained in studies of similar complexes using TRIR have been powerful and informative (§1.4.1 and §2.3).

3.5.1. UV/visible Absorption and Luminescence Spectroscopy

UV/vis Absorption Spectroscopy

UV/visible absorption spectra of RuCN-BL\textsuperscript{1}-Rubpyamide have been measured in CH\textsubscript{3}CN and D\textsubscript{2}O (see Figure 3.5.1.1). The spectral profile and peak positions correspond to an approximate superposition of the absorption spectra of the model complexes RuCNdmb and Rubpyamide-BL\textsuperscript{1}, which supports the notion there is no electronic communication between the two metal centres in their electronic ground states. The peak positions and assignments (by comparison with model complexes in §1.4.1.1 and §2.3.1) are summarised in Table 3.5.1.1.

It is apparent from visual inspection of the spectra that two of the longer wavelength features in the spectrum are relatively solvent-invariant (<10 nm difference between CH\textsubscript{3}CN and D\textsubscript{2}O) and these can be assigned to Ru \( \rightarrow \) (4,4'-CO(NEt\textsubscript{2})\textsubscript{2}bpy) \textsuperscript{1}MLCT
transitions. The weaker remaining feature varies over a range of ca. 85 nm and exhibits negative solvatochromism as is expected for ruthenium (II) cyanide complexes and is assigned to a Ru → (5,5'-((CH$_3$)$_2$)bpy $^1$MLCT transition.

![Absorbance (Normalised)](image)

**Figure 3.5.1.1:** The UV/visible absorption spectra of RuCN-BL$^1$-Rubpyamide in D$_2$O (black) and CH$_3$CN (red). The absorbances are normalised to the 300 nm feature. Arrows denote the position of the strongly solvatochromic peak in the two solvents.

<table>
<thead>
<tr>
<th>Solvent</th>
<th>$\lambda_{max}$ / nm</th>
</tr>
</thead>
<tbody>
<tr>
<td>D$_2$O</td>
<td>&lt; 300  369  403  439  478</td>
</tr>
<tr>
<td>CH$_3$CN</td>
<td>&lt; 300  367  432  468  505</td>
</tr>
</tbody>
</table>

<table>
<thead>
<tr>
<th>Assignment</th>
<th>Terminus</th>
<th>$^1$MC</th>
<th>$^1$MLCT</th>
<th>$^1$MLCT</th>
<th>$^1$MLCT</th>
</tr>
</thead>
<tbody>
<tr>
<td>IL</td>
<td>$\pi$-$\pi^*$</td>
<td>RuCN</td>
<td>Rubpyamide</td>
<td>RuCN</td>
<td></td>
</tr>
</tbody>
</table>

**Table 3.5.1.1:** Table of peak positions and assignments for UV/visible absorption spectra of RuCN-BL$^1$-Rubpyamide in D$_2$O and CH$_3$CN solutions.
Luminescence Spectroscopy

Luminescence spectroscopy was used to study RuCN-BL₁-Rubpyamide in D₂O and CH₃CN and are the spectra measured are shown in Figure 3.5.1.2 and 3.5.1.3 respectively. Figure 3.5.1.4 compares emission in the two solvents.

![Emission Spectra](image)

**Figure 3.5.1.2: Uncorrected emission spectra of RuCN-BL₁-Rubpyamide in D₂O, with λ_ex = 476 nm (black), λ_ex = 429 nm (red) and λ_ex = 394 nm (green). The spectra were recorded under identical conditions.**

The complex is luminescent giving a broad, featureless spectrum with λ_max (H₂O) = 656 nm and λ_max (CH₃CN) = 646 nm. There is a substantial difference in the emission intensity measured in the two solvents; the luminescence is much weaker in CH₃CN than in D₂O. The emission wavelengths and quantum yields measured are presented in Table 3.5.1.2.
Figure 3.5.1.3: Uncorrected emission spectra of RuCN-BL\textsuperscript{1}-Rubpyamide in CH\textsubscript{3}CN, with $\lambda_{\text{ex}} = 480$ nm (black), $\lambda_{\text{ex}} = 433$ nm (red) and $\lambda_{\text{ex}} = 397$ nm (green).

The spectra were recorded under identical conditions.

<table>
<thead>
<tr>
<th>D\textsubscript{2}O</th>
<th>CH\textsubscript{3}CN</th>
</tr>
</thead>
<tbody>
<tr>
<td>$\lambda_{\text{ex}}$/ nm</td>
<td>$\lambda_{\text{em}}$/ nm</td>
</tr>
<tr>
<td>476</td>
<td>656</td>
</tr>
<tr>
<td>429</td>
<td>656</td>
</tr>
<tr>
<td>394</td>
<td>656</td>
</tr>
</tbody>
</table>

Table 3.5.1.2: Table of peak positions and quantum yields from luminescence spectroscopy of RuCN-BL\textsuperscript{1}-Rubpyamide in D\textsubscript{2}O and CH\textsubscript{3}CN solution.

This is consistent with the emissive states being $^3$MLCT in nature, with the vast majority of the detected luminescence being derived from the Rubpyamide terminus (Rubpyamide-BL\textsuperscript{1} emits at ca. 650 nm).
Figure 3.5.1.4: Uncorrected emission spectra of RuCN-BL\textsuperscript{1}-Rubpyamide in D\textsubscript{2}O (black) and CH\textsubscript{3}CN (red), with \( \lambda_{ex} (D\textsubscript{2}O) = 394 \text{ nm} \) and \( \lambda_{ex} (CH\textsubscript{3}CN)= 397 \text{ nm} \). The spectra were recorded under identical conditions.

These results confirm that the photoprocesses occurring vary significantly in different solvents. We can rationalise these results in the framework of our photophysical model for these complexes, based on the work of Simpson \textit{et al.} (§3.1).\textsuperscript{3} In CH\textsubscript{3}CN the \( ^3\text{MLCT} \) excited state of the Rubpyamide terminus is the higher in energy of the two and thus luminescence is partially quenched by energy transfer to the RuCN\textsubscript{dmb} terminus, which only emits weakly. There is a \textit{ca.} 90% reduction in quantum yield from that of Rubpyamide-BL\textsuperscript{2} in CH\textsubscript{3}CN. This partial quenching offers us further insight into the efficiency of the PEnT process in this system. In D\textsubscript{2}O the \( ^3\text{MLCT} \) excited state of the Rubpyamide terminus is the lower in energy and thus its luminescence remains unquenched at an intensity comparable to that of Rubpyamide-BL\textsuperscript{2}; there is less than 10% difference between the values obtained for the two complexes. The luminescence from the RuCN\textsubscript{dmb} terminus is particularly difficult to detect in this case as it occurs at a wavelength coincident with the intense luminescence from the Rubpyamide terminus.
TCSPC Measurements

A kinetic analysis of the luminescence of RuCN-BL¹-Rubpyamide was performed using 405 nm excitation and registration of emission near the peak maximum (640 nm and 629 nm in D₂O and CH₃CN respectively). It was found that the traces obtained could not be fitted using a simple exponential function; indeed at least 4 components with time constants varying by orders of magnitude were required to obtain acceptable fits to the experimental data ($\chi^2 < 1.05$). The results are summarised in Table 3.5.1.2 and the attempts to fit the experimental data to between 1 and 4 components in D₂O and CH₃CN is displayed in Figures 3.5.1.5 and 3.5.1.6 respectively. From these it is clear that at least 4 components are required for satisfactory fits. The processes responsible for this complicated time dependence are not immediately clear but $\tau_1$ is thought to correspond to the Instrument Response Function (IRF).

<table>
<thead>
<tr>
<th>Component</th>
<th>Time constant / ns</th>
<th>% Contribution</th>
<th>Time constant / ns</th>
<th>% Contribution</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>D₂O ($\chi^2 = 0.980$)</td>
<td>CH₃CN ($\chi^2 = 1.024$)</td>
<td></td>
<td></td>
</tr>
<tr>
<td>$\tau_1$</td>
<td>5.42 (±0.190)</td>
<td>8.07</td>
<td>1.93 (±0. 163)</td>
<td>1.95</td>
</tr>
<tr>
<td>$\tau_2$</td>
<td>16.8 (±2.82)</td>
<td>39.2</td>
<td>7.57 (±0.241)</td>
<td>8.77</td>
</tr>
<tr>
<td>$\tau_3$</td>
<td>60.4 (±4.3)</td>
<td>5.81</td>
<td>81.6 (±7.64)</td>
<td>5.39</td>
</tr>
<tr>
<td>$\tau_4$</td>
<td>487.1 (±3.15)</td>
<td>47.0</td>
<td>309.61 (±19.3)</td>
<td>83.89</td>
</tr>
</tbody>
</table>

Table 3.5.1.2: Table of exponential fit components of luminescence spectra of RuCN-BL¹-Rubpyamide in D₂O and CH₃CN solutions.

The kinetic data obtained is clearly polyexponential in form, although it is important to note that $\tau_1$ in both solvents may arise from the IRF. These results suggest that there are a number of photophysical processes occurring in solution on the nanosecond timescale. We have used TRIR spectroscopy to study these processes in more detail (§3.5.3).
Figure 3.5.1.5: Fits to lifetime data (with residuals) for RuCN-BL'-Rubpyamide in D_2O. Top left: 1 exponent; top right: 2 exponents; bottom left: 3 exponents, bottom right: 4 exponents.
Figure 3.5.1.6: Fits to lifetime data (with residuals) for RuCN-BL\textsuperscript{1}-Rubpyamide in CH\textsubscript{3}CN. Top left: 1 exponent; top right: 2 exponents; bottom left: 3 exponents, bottom right: 4 exponents.
We can rationalise these results in the framework of our model for the complex. In CH₃CN the 3MLCT excited state of the Rubpyamide terminus is the higher in energy of the two and thus luminescence is partially (ca. 90%) quenched by energy transfer to the RuCN_dmb terminus, which itself emits weakly with barely detectable luminescence. This partial quenching may offer us further insight into the efficiency of the PEnT process in this system. In D₂O the 3MLCT excited state of the Rubpyamide terminus is the lower in energy and thus its luminescence is expected to remain unquenched at an intensity comparable to that of Rubpyamide-BL¹. This is not the case, with the quantum yields differing by an order of magnitude. The luminescence from the RuCN_dmb terminus is extremely difficult to detect in this case as it occurs at a wavelength coincident with the intense luminescence from the Rubpyamide terminus. The heavily overlapped absorption profiles of the two termini resulting in unselective excitation of either terminus by incident photons complicates the situation further.

A possible reason for the complexity of the kinetic data obtained for RuCN-BL¹-Rubpyamide is the conformationally flexible nature of the bridging ligand BL¹. There are axes of rotation about the CH₂CH₂ linker between the two bipyridines which allow the complex to rotate freely in a room temperature solution. This may lead to an ensemble of supramolecular conformers being present in solution at any one time, each of which possessing a different interchromophoric (i.e. donor-acceptor) distance (r, §1.1.3). This is commonly observed in the luminescence of biological systems.²¹

The Förster equation expresses the sensitivity of the rate of through-space energy transfer in mathematical terms (§1.1.3). In the equation, the rate is proportional to the sixth power of the donor-acceptor distance, meaning that even very small changes in intermetallic distances will lead to profound differences in the kinetics of the photoinduced processes occurring in solution, and changes on the magnitude of molecular rotation about the CH₂CH₂ axis of BL¹ will lead to a wide range of PEnT rates. This conformational flexibility may also help to explain why the quenching of
luminescence in CH₃CN is only partial – Förster PEnT requires good spatial overlap for efficient energy transfer and if the supramolecular system spends a great deal of its time in geometries which are not favourable for PEnT then energy transfer may have to compete with other radiative and non-radiative excited state decay processes.

Studies of Cyanide-Metal Ion Interactions
The effect of metal ion binding to the cyanide ligands of RuCN-BL¹-Rubpyamide was also studied using luminescence spectroscopy. The variation of emission wavelength and intensity was studied as a function of Ba²⁺ concentration in CH₃CN solution. As [Ba²⁺] was increased the luminescence intensity increased substantially (from $\phi_{\text{init}} = 0.001$ to $\phi_{\text{max}} = 0.022$). The spectra recorded are presented in Figure 3.5.1.6.

The physical mechanism responsible for such a dramatic change in emission properties is analogous to that reported by Simpson et al. when exploring the solvent dependence of RuCN-BL-Rubpy in DMSO:H₂O solvent mixtures (§3.1, Figure 3.1.1.4). In that study, as the A. N. of the solvent mixture increased through titration of H₂O into the solution, the intensity of measured emission increased as the ³MLCT energy of the RuCN terminus of the bimetallic complex was raised through SSDA interactions. However in this case the modulation of RuCN ³MLCT energy occurs through control of [Ba²⁺]. The process of Ba²⁺ binding to RuCN-BL¹-Rubpyamide leads to the reversal of the location of the lowest energy excited state from the RuCN terminus to the Rubpyamide terminus. The crucial consequence of this is that as [Ba²⁺] is increased, Rubpyamide $\rightarrow$ RuCN PEnT (which is ostensibly a major luminescence quenching process) becomes thermodynamically unfavourable and the luminescence intensity increases substantially. Addition of a macrocyclic Ba²⁺ sequestration agent (dibenzo-18-crown-6) demonstrates the near total reversibility of this process as the luminescence returns to a level similar to that measured in the absence of Ba²⁺.

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This study demonstrates that the excited state energy of the RuCN terminus is amenable to reversible modulation by other external factors such as metal ion binding in addition to solvent composition. The experiments discussed above form part of a study presented in a forthcoming publication which explores this effect in more detail.\textsuperscript{22}

**Figure 3.5.1.6:** Changes in luminescence spectra of RuCN-BL\textsuperscript{1}-Rubpyamide in CH\textsubscript{3}CN solution upon addition of Ba\textsuperscript{2+} ions (solid lines) and dibenzo-18-crown-6 (dashed line). Relative concentrations corresponding to the line colours are denoted in the legend.

3.5.2. Cyclic Voltammetry

RuCN-BL\textsuperscript{1}-Rubpyamide has been studied using cyclic voltammetry in CH\textsubscript{3}CN (see Appendix §3.9 for the voltammogram) and the results are presented in Table 3.5.2.1. Within the potential window of CH\textsubscript{3}CN (ca. +2 V to -2 V) there is one reversible oxidation, a large irreversible feature at positive potential (due to oxidation of the amide groups) and three reversible reduction couples arising from the bpy-based
ligands. The potentials for the oxidation and the 1st and 2nd reductions are very similar to those reported for Rubpyamide-BL¹ and Rubpyamide-BL² (§3.3.2, 3.4.2).

<table>
<thead>
<tr>
<th>$E_{\text{w}}/V$ (vs. Fe/Fe⁺)</th>
<th>Reversibility</th>
<th>Assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>+0.90</td>
<td>Reversible</td>
<td>Ru¹/II/III</td>
</tr>
<tr>
<td>$(E_p) +0.71$</td>
<td>Irreversible</td>
<td>$(4,4'-(CONEt₂)₂bpy)^0/\pm$</td>
</tr>
<tr>
<td>-1.58</td>
<td>Reversible</td>
<td>$(4,4'-(CONEt₂)₂bpy)^0/\pm$</td>
</tr>
<tr>
<td>-1.78</td>
<td>Reversible</td>
<td>$(4,4'-(CONEt₂)₂bpy)^0/\pm$</td>
</tr>
<tr>
<td>&lt;= -2.00</td>
<td>Reversible</td>
<td>(BL¹)^0/\pm</td>
</tr>
</tbody>
</table>

Table 3.5.2.1: Table of potentials for redox couples and assignments for the cyclic voltammogram of RuCN-BL¹-Rubpyamide in CH₃CN solution.

These results are expected from a bimetallic complex with no intermetallic interactions in the ground state which contains a Rubpyamide-BL¹ unit (§3.1, 3.3.2, 3.4.2), as the Rubpyamide redox couples fall within the solvent window and are in very close agreement with those measured for the model complex. The Ru¹/II/III couple of the RuCN terminus is expected to be present but has not been observed.

3.5.3. FTIR and Time-Resolved Infrared Spectroscopy

Luminescence spectroscopy has provided some indirect evidence for PEnT between the two termini of RuCN-BL¹-Rubpyamide (vide supra), through the quenching of emission from the Rubpyamide terminus in low A. N. solvent environments. TRIR spectroscopy has been used to monitor this process in order to obtain direct evidence for its occurrence. The system has two sets of IR active ‘reporters’, one on each terminus. Both can provide important and complementary information and it is crucial to consider them together when analysing the data.

It is important to remember that due to the very similar nature of the MLCT chromophores appended to the two metal centres and the broadness of $^1$MLCT
absorption peaks, there is often a case of unselective excitation, *i.e.* either metal centre may be excited by the incident laser light. This adds another degree of complexity to the study, as there may be more than one excited state species present in solution and as a result *bona fide* PENT processes must be disentangled from the other primary photophysical processes occurring in order to fully understand the spectra recorded.

**Deuterium Oxide Solution**

In D₂O, SSDA interactions elevate the energy of the lowest energy RuCN $^3$MLCT state above the corresponding state in the Rubpyamide terminus. Therefore RuCN → Rubpyamide PENT is expected to take place.

The FTIR spectrum of RuCN-BL¹-Rubpyamide in D₂O solution in the v(CN) and amide v(CO) regions has been measured (see Appendix §3.9 for spectra). In the cyanide region there is a peak profile characteristic of the tetracyanoruthenate (II) unit with 3 resolved peak maxima at 2091, 2050 and 2030 cm⁻¹. In the amide region there is a single peak present at 1612 cm⁻¹. These results are closely comparable with the model complexes RuCNdmb and Rubpyamide-BL¹ in D₂O solution (§3.2.5, §3.3.5).

Picosecond and nanosecond TRIR spectra (with laser excitation at 400 and 355 nm respectively) of RuCN-BL¹-Rubpyamide were recorded in D₂O solution, monitoring both regions in the IR where reporter groups are present.

**ps-TRIR Spectroscopy**

Figure 3.5.3.1 shows the ps-TRIR spectra obtained for the RuCN and Rubpyamide termini in D₂O.

**RuCN ("Donor") terminus**

The spectral profile obtained following 400 nm excitation resembles that of RuCNdmb in D₂O (§3.2.5), with several overlapped bleaches centred between 2030 and 2100 cm⁻¹ and at least two transient bands detected at higher wavenumber between 2090 and 2125 cm⁻¹. There appear to be changes in the profile on the
picosecond timescale, particularly in the region where the high energy bleach is overlapping with the transient. Additional fine structure is observed in the spectrum in comparison with the corresponding spectra in CH$_3$CN. A possible reason for these additional features is vibrational cooling; however the kinetic data does not definitively exhibit characteristic narrowing and blue-shifting of transient bands on the appropriate timescale ($\tau \sim 10$ ps, \textit{vide infra}) and the solvent-dependent nature of this phenomenon suggests that it is related to the solvatochromic effect (§1.2.3). An alternative hypothesis of a multi-step photoinduced process (ostensibly PEnT) involving the formation and subsequent interconversion of two discrete transient species (plausible in D$_2$O only) is posited and discussed in detail in the Discussion section below (§3.5.5).

Over the time range of the experiment ($\Delta t = 1$-1000 ps), the majority of the v(CN) bleach and transient signals recover, indicating that there are photophysical processes occurring which facilitate the quenching of the RuCN$_{dmb}$ $^3$MLCT excited state. At $\Delta t = 1$ ns only a weak residual signal remains, indicating that the majority of RuCN$_{dmb}$ excited states formed in solution have been quenched.

**Rubpyamide ("Acceptor") terminus**

The spectra obtained for the Rubpyamide terminus in D$_2$O resemble the profile observed for the precursor complex Rubpyamide-BL$^1$ (§3.5.5), with an amide v(CO) bleach centred at 1615 cm$^{-1}$ and a transient to lower wavenumber centred at 1591 cm$^{-1}$. This is consistent with the anticipated $^3$MLCT lowest excited state for this terminus. The spectra show much less development over the timescale of the experiment than for the RuCN terminus, indicating that ultrafast processes are not contributing to excited state quenching to the same extent as in the "donor" terminus. Indeed at $\Delta t = 1$ ns there is still a strong signal remaining, indicating that excited state decay is taking place on longer timescales.
Figure 3.5.3.1: ps-TRIR spectra of RuCN-BL'-Rubpyamide in D$_2$O solution in the v(CN) and amide v(CO) region. The fine structure observed in these spectra is discussed in more detail in §3.5.5.
This initial result is consistent with our model, which predicts that the RuCN donor terminus will be quenched by efficient PEnT and have a short-lived excited state, whilst the Rubpyamide acceptor terminus will possess a long-lived excited state owing to it being the ultimate excited state being formed.

**ps-TRIR Kinetics**

Figures 3.5.3.2 and 3.5.3.3 show the ps-TRIR kinetics (for bleach and transient where possible) obtained for the RuCN and Rubpyamide termini in different time domains. Figure 3.5.3.2 shows $\Delta t = 0$-50 ps and 3.5.3.3 shows $\Delta t = 0$-1 ns. The reader is reminded that a hypothesis detailing a multi-step PEnT process in RuCN-BL$^1$-Rubpyamide is discussed in detail in §3.5.5 and includes a rationalisation of the apparently anomalous single-pixel kinetic trace presented below for the RuCN terminus in D$_2$O (Figure 3.5.3.2).

$\Delta t = 0$-50 ps

This kinetics recorded during this period are interesting as there is clear evidence for the growth in magnitude of both the bleach and the transient on the Rubpyamide terminus over the first ca. 30 ps. This is accompanied by a depletion in the magnitude of both the bleach and the transient on the RuCN terminus over the same timescale. The observation of this effect in both bleach and transient kinetic data - taken from pixels far apart from each other to minimise overlap effects - rules out the possibility of vibrational cooling being responsible for this phenomenon. If it were vibrational cooling, we would only observe this effect in the transient data. The signal intensity is such that the time constants obtained are not precise, but they are all within the region of 5-10 ps, which would correspond to a time constant of 8 (±3) ps for efficient and rapid PEnT in this supramolecular system.

$\Delta t = 0$-1 ns

In this timescale the processes described above are observed in the early part of the trace, and are followed by signal decay on both termini, indicating that excited state decay by routes other than PEnT has become the major process occurring. As was observed in the luminescence spectroscopy of these complexes, these processes appear to be multi-exponent. This may be due to the offsetting of
excited state decay processes with PEnT-related grow-in of signal - or the opposite situation where there are multiple competing decay processes occurring at varying rates. The time constant for the major decay component in each case is in the region of 300-600 ps.

**ns-TRIR Spectroscopy and Kinetics**

As observed by luminescence spectroscopy, the **Rubpyamide** terminus has a long-lived excited state in D$_2$O which can be monitored on the nanosecond timescale. Figures 3.5.3.4 and 3.5.3.5 show the spectroscopy and kinetics measured at $\Delta t = 1$-500 ns. Attempts to obtain ns-TRIR spectra of the **RuCNdmb** terminus were not successful, conceivably due to its short excited state lifetime.

The ns-TRIR spectral profile measured for **Rubpyamide** following excitation at 355 nm is analogous to the ps-TRIR spectra shown above, with the parent bleach centred at 1616 cm$^{-1}$ and a transient band centred at 1591 cm$^{-1}$. The signals detected were weak, implying that much of the excited state had decayed on the picosecond timescale, although the signal intensities over the two timescales are unfortunately not directly comparable due to different laser excitation wavelengths, data collection times and sample concentrations.

Analysis of the kinetics at the bleach and transient peak maxima give multiexponent fits (the data were fitted to at least 3), with components having time constants of 5.5 (±3), 25 (±9) and 550 (±200) ns. This is consistent with the luminescence data presented in §3.5.3.1 and confirms that there are multiple photophysical processes occurring in this system over the timescales investigated.

**Acetonitrile Solution**

In CH$_3$CN SSDA interactions stabilise the energy of the **RuCN** $^3$MLCT state so that it is lower than the corresponding state in the **Rubpyamide** terminus. Therefore **Rubpyamide $\rightarrow$ RuCN PEnT** is expected to take place.
Figure 3.5.3.2: Single pixel ps-TRIR kinetics of RuCN-BL$^1$-Rubpyamide in D$_2$O solution for the RuCN (transient at 2101 cm$^{-1}$, top left and bleach at 2062 cm$^{-1}$, bottom left) and Rubpyamide (transient at 1584 cm$^{-1}$, top right and bleach at 1613 cm$^{-1}$, bottom right) termini for $\Delta t =$ 0-50 ps. The anomalous trace obtained for the RuCN transient is discussed in detail in §3.5.5.
Figure 3.5.3.3: ps-TRIR kinetics of RuCN-BL$^1$-Rubpyamide in D$_2$O solution for the RuCN (transient at 2101 cm$^{-1}$, top left and bleach at 2062 cm$^{-1}$, bottom left) and Rubpyamide (transient at 1584 cm$^{-1}$, top right and bleach at 1613 cm$^{-1}$, bottom right) termini for $\Delta t$ = 0-1 ns.
The FTIR spectrum of RuCN-BL\textsuperscript{1}-Rubpyamide in CH\textsubscript{3}CN solution in the v(CN) and amide v(CO) region has been measured (see Appendix §3.9 for spectra). In the cyanide region there is a peak profile characteristic of the tetracyanoruthenate (II) unit with 3 resolved peak maxima at 2093, 2069 and 2044 cm\textsuperscript{-1}. In the amide region there is a single peak present at 1638 cm\textsuperscript{-1}. These results are closely comparable with the constituent complexes RuCN\textsubscript{dmb} and Rubpyamide-BL\textsuperscript{1} in CH\textsubscript{3}CN solution (§3.2.5, §3.3.5).

Picosecond and nanosecond-TRIR spectroscopy (with laser excitation at 400 and 355 nm respectively) were performed on RuCN-BL\textsuperscript{1}-Rubpyamide in CH\textsubscript{3}CN solution, monitoring both regions in the IR where reporter groups are present.

**ps-TRIR Spectroscopy**

Figure 3.5.3.6 shows the ps-TRIR spectra obtained for the RuCN and Rubpyamide termini in CH\textsubscript{3}CN solution.

**RuCN ("Acceptor") terminus**

The spectral profile obtained following 400 nm excitation resembles that of RuCN\textsubscript{dmb} in CH\textsubscript{3}CN (§3.2.5), with several overlapped bleaches centred between 2050 and 2115 cm\textsuperscript{-1} and transient bands detected at higher wavenumber between 2100 and 2140 cm\textsuperscript{-1}. There appear to be changes in the profile on the picosecond timescale, particularly in the region where the high energy bleach is overlapping with the transient. It is plausible that vibrational cooling may be responsible for these changes but no definitive spectral evidence has been observed. Over the time range of the experiment (1-1000 ps), the majority of the v(CN) bleach and transient signals recover, albeit to a lesser extent than was observed in D\textsubscript{2}O. At $\Delta t = 1$ ns there is still a significant residual signal remaining, indicating that the quenching of RuCN\textsubscript{dmb} excited states formed in CH\textsubscript{3}CN is much less efficient than in D\textsubscript{2}O.
Figures 3.5.3.4 and 3.5.3.5: ns-TRIR spectra (left) and kinetics (transient at 1589 cm\(^{-1}\), top right and bleach at 1611 cm\(^{-1}\), bottom right) of RuCN-BL\(^{I}\)-Rubpyamide in D\(_2\)O solution for the Rubpyamide terminus for \(\Delta t = 0\) - 300 ns.
Rubpyamide ("Donor") terminus
The spectra obtained for the Rubpyamide terminus in CH$_3$CN resemble the profile observed for the precursor complex Rubpyamide-BL$^1$ (§3.3.5), with an amide v(CO) bleach centred at 1637 cm$^{-1}$ and a transient to lower wavenumber centred at 1620 cm$^{-1}$. This is consistent with a $^3$MLCT excited state for this terminus. In contrast to the observation in D$_2$O, there is a much greater extent of transient decay and parent recovery, such that over 95% of the integrated absorbance of the bleach has decayed by $\Delta t = 1$ ns. The spectra have poor signal-to-noise ratios and were collected with great difficulty, perhaps due to a combination of low oscillator strength of the v(CO) amide group and rapid quenching of the Rubpyamide terminus excited state leading to low concentrations of this species being present in solution.

This result is consistent with our model, which predicts that the Rubpyamide donor terminus will be quenched by efficient PEnT and have a short-lived excited state, whilst the RuCN acceptor terminus will possess a long-lived excited state owing to it being the ultimate excited state being formed. The slight differences in the proportion of signal intensity remaining at $\Delta t = 1$ ns for the two termini between D$_2$O and CH$_3$CN can be understood by taking into account the very different natural excited state lifetimes of the two units (as measured in the model complexes RuCNdmb and Rubpyamide-BL$^1$, §3.2.5, §3.3.5) and indeed the very significant effect that solvent composition has on the excited state properties of RuCNdmb in particular.

ps-TRIR Kinetics
Figures 3.5.3.7 and 3.5.3.8 show the ps-TRIR kinetics (for bleach and transient where possible) obtained for the RuCN and Rubpyamide termini in different time domains in CH$_3$CN solution. Figures 3.5.3.7 shows the $\Delta t = 0$-50 ps time range and 3.5.3.8 shows $\Delta t = 0$-1 ns.

$\Delta t = 0$-50 ps
This kinetics recorded during this period are less definitive than those presented for D$_2$O, as there is clear evidence for the growth in magnitude of both the bleach and the transient on the RuCN terminus over the first ca. 30 ps, but unfortunately
no data of comparable clarity for the **Rubpyamide** terminus. For the **RuCN** terminus, which as the acceptor provides the tell-tale grow-in of absorbance on both bleach and transient bands, the signal intensity is such that precise time constants could not be meaningfully obtained. However it is of note that all four of the termini within the region of 9-15 ps. This would correspond to an approximate time constant of 12 (±3) ps for efficient and rapid PEnT in this supramolecular system. Analysis of the donor terminus data provides few conclusions as no clear processes (grow-in or decay) are observed over the Δt = 0-50 ps timescale. A possible reason for this is the offsetting of decay processes in the **Rubpyamide** terminus by peak sharpening due to vibrational cooling of the ν(CO) bands.

\[ Δt = 0-1 \text{ ns} \]

As was observed in D_2O solution, during this time domain the processes described above are observed at early times following excitation, and these are followed by signal decay on both termini, indicating that the excited state decay by mechanisms other than PEnT becomes the major process occurring. These processes appear to be multi-exponent, with fits utilising 1 or 2 components most closely representing the data. This may be due to the offsetting of excited state decay processes with PEnT-related grow-in of signal – or the converse situation where there are multiple competing decay processes occurring at varying rates. There appears to be little consistency in the time constants determined for the fits to the experimental data, with time constants varying between 20 and 500 ps for the major decay component.

**ns-TRIR Spectroscopy and Kinetics**

In CH₃CN our photophysical model predicts that the **RuCN** terminus will have a long-lived \(^3\)MLCT excited state which can be monitored on the nanosecond timescale. Figures 3.5.3.9 and 3.5.3.10 show the spectroscopy and kinetics measured at Δt = 1-500 ns. TRIR spectra of the **Rubpyamide** terminus in CH₃CN solution on the nanosecond timescale showed no transient features, conceivably due to the short excited state lifetime of the \(^3\)MLCT excited state located on the **Rubpyamide** terminus.
Figure 3.5.3.6: ps-TRIR spectra of RuCN-BL¹-Rubpyamide in CH₃CN solution in the v(CN) and amide v(CO) region.
The ns-TRIR spectroscopy of the RuCN terminus following excitation at 355 nm is analogous to the ps-TRIR spectroscopy shown above, with several overlapped bleaches centred between 2050 and 2085 cm\(^{-1}\) and a transient band at higher wavenumber centred at 2095 cm\(^{-1}\). The signals detected were weak, implying that most of the excited state had decayed on the picosecond timescale, although the signal intensities over the two experiments are unfortunately not directly comparable due to different laser excitation wavelengths, data collection times and sample concentrations.

Analysis of the kinetics at the bleach and transient peak maxima give biexponential fits, with components having time constants of 3.0 (±1.5) and 65 (±12) ns. This unfortunately cannot be compared with the corresponding luminescence experiment as the weak RuCN luminescence was undetectable alongside the much more intense emission from the Rubpyamide terminus. This confirms that there are multiple photophysical processes occurring in this system over the timescales investigated.

**Inter or Intramolecular Processes?**

Due to the electrostatic attractions between the RuCN and Rubpyamide termini in this system, it is conceivable that intermolecular “ion pairs” or aggregates may form in solution, and there is the possibility that *intermolecular* photophysical processes such as PEnT may take place (*i.e.* between more than one supramolecular complex). To investigate the nature of these processes, mixtures of the model complexes RuCN\(\text{dmb}\) and Rubpyamide-\(\text{BL}^1\) were studied using TRIR spectroscopy on the picosecond and nanosecond timescales in CH\(_3\)CN (not shown). No decay or growth processes were observed on the picosecond timescale on either terminus, and on the nanosecond timescale the only process observed coincides well with the decay lifetimes of the model complexes in isolation (§3.2.5, 3.3.5). This provides evidence that intermolecular processes (including PEnT) are not a major consideration in these complexes and the photophysical processes observed in RuCN-BL\(^1\)-Rubpyamide are *intramolecular* in nature.
Figure 3.5.3.7: ps-TRIR kinetics of RuCN-BL$^+$-Rubpyamide in CH$_3$CN solution for the RuCN (transient at 2114 cm$^{-1}$, top left and bleach at 2068 cm$^{-1}$, bottom left) and Rubpyamide (transient at 1616 cm$^{-1}$, top right and bleach at 1637 cm$^{-1}$, bottom right) termini for $\Delta t = 0-50$ ps.
Figure 3.5.3.8: ps-TRIR kinetics of RuCN-BL \textsuperscript{1} -Rubpyamide in CH\textsubscript{3}CN solution for the RuCN (transient at 2114 cm\textsuperscript{-1}, top left and bleach at 2068 cm\textsuperscript{-1}, bottom left) and Rubpyamide (transient at 1616 cm\textsuperscript{-1}, top right and bleach at 1637 cm\textsuperscript{-1}, bottom right) termini for \( At = 0-1 \) ns.
Figures 3.5.3.9 and 3.5.3.10: ns-TRIR spectra (left) and kinetics (transient at 2085 cm⁻¹, top right and bleach at 2070 cm⁻¹, bottom right) of RuCN-BL'Rubpyamide in CH₃CN solution for the RuCN terminus for Δt = 0-600 ns.
3.5.4. Computational Studies and Conformational Analysis

As discussed throughout this section, RuCN-BL\textsuperscript{1}-Rubpyamide has a considerable amount of conformational freedom due to the ability to rotate freely around the CH\textsubscript{2}CH\textsubscript{2} linkage of the BL\textsuperscript{1} ligand. Molecular mechanics calculations have been performed using the SPARTAN02 computational suite to gain insight into the effect that this conformational freedom has on the parameters that govern the rate and efficiency of PEnT.

Four important cases of conformational behaviour have been isolated, as shown in Figure 3.5.4.1. These consist of two 'closed' structures where the two metal centres are in close proximity (although their ligands prevent closer approach) and two 'open' conformers in which the BL\textsuperscript{1} ligand has unfolded to form a near linear chain. The relative positions of the metal centres in these two pairs of conformers can be thought of as syn (i and iii) or anti (ii and iv) to each other. The intermetallic distances in these cases range from 5.4 Å in conformer (i) to 12.7 Å in conformer (iv). Therefore it is likely that an ensemble of conformers are present in solution with intermetallic distance $5.41 < r < 12.7 \text{ Å}$.

It is important to consider the electrostatics of RuCN-BL\textsuperscript{1}-Rubpyamide when estimating the relative proportions of these three conformers present in solution. As the RuCN terminus is doubly anionic and the Rubpyamide terminus is doubly cationic, there is a strong likelihood that ion pairing will take place to some extent, perhaps skewing the conformer distribution towards the 'closed' structure of (i). As this conformer exhibits the shortest intermetallic distance it would also be optimal for efficient PEnT.
Figure 3.5.4.1: Four important conformers of RuCN-BL¹-Rubpyamide calculated in the gas phase using SPARTAN's molecular mechanics suite: i) and ii) are syn and anti folded conformers of closed structures; iii) and iv) are syn and anti conformers of open structures. Amide groups and H atoms omitted for clarity.
3.5.5. RuCN-BL\textsuperscript{1}-Rubpyamide: Discussion and Closing Points

The bimetallic complex RuCN-BL\textsuperscript{1}-Rubpyamide has been studied using UV/visible absorption and luminescence spectroscopy, cyclic voltammetry and TRIR spectroscopy in D\textsubscript{2}O and CH\textsubscript{3}CN solutions. In addition to solvatochromic studies a brief investigation into metallochromism was also performed by way of a luminescence titration with Ba\textsuperscript{2+} ions.

UV/visible spectroscopy (§3.5.1) and cyclic voltammetry (§3.5.2) of RuCN-BL\textsuperscript{1}-Rubpyamide have shown that the ground state properties of the system are essentially a superposition of the constituent monometallic fragments most closely represented by the model complexes RuCN\textsubscript{dmb} and Rubpyamide-BL\textsuperscript{1}. This supports our supposition that the saturated spacer connecting the two bpy ligands in the bridging ligand BL\textsuperscript{1} prevents electronic communication between the two termini. UV/vis absorption spectroscopy has observed a measurable difference in the absorption profiles of RuCN-BL\textsuperscript{1}-Rubpyamide in CH\textsubscript{3}CN and D\textsubscript{2}O which is more subtle than observed in the model complex RuCN\textsubscript{dmb}. This is presumably due to the presence of one terminus exhibiting strong negative solvatochromism (RuCN) and one which is essentially solvent-independent (Rubpyamide). These results are consistent with those obtained by Simpson \textit{et al.} in the study of RuCN-BL-Rubpy (§3.1).\textsuperscript{3}

Luminescence spectroscopy of RuCN-BL\textsuperscript{1}-Rubpyamide has shown that the spectral profile is dominated by emission arising from the Rubpyamide terminus. The considerable solvent-dependence of the emission quantum yield and multi-exponent nature of the lifetime has been studied with experiments in CH\textsubscript{3}CN and D\textsubscript{2}O. The emission is quenched in CH\textsubscript{3}CN relative to D\textsubscript{2}O, consistent with Rubpyamide → RuCN PEnT being thermodynamically favourable in low A. N. solvents such as CH\textsubscript{3}CN. The presence of multiexponential kinetics may be related to the conformational freedom possessed by the system (§3.5.4, \textit{vide infra}). As with RuCN-BL-Rubpy (§3.1), the emission is considerably more intense in high A. N. solvents.
(\(\varphi_{\text{D}_2\text{O}} > \varphi_{\text{CH}_3\text{CN}}\)) which may be rationalised using an analogous model (§3.5.1, Figure 3.1.1.4). Also in keeping with literature results for RuCN-BL-Rubpy (§3.1) emission from the RuCN terminus was not detected, presumably owing to its low quantum yield and/or being obscured by emission from the much more intense Rubpyamide terminus.

The metallochromic properties of RuCN-BL\(^1\)-Rubpyamide have been demonstrated using luminescence spectroscopy\(^{23}\) and provide an interesting avenue for future work in the areas of metal ion sensing and designing photophysical switches which can respond to multiple independent external stimuli (solvent A. N. and \([M^+]\)); this may be a requirement in the design and preparation of future photochemical device components.\(^{24}\)

TRIR spectroscopy has been used to directly monitor the processes occurring following photoexcitation of RuCN-BL\(^1\)-Rubpyamide on the picosecond and nanosecond timescales. In agreement with the model presented in §3.5.3, the results presented in the same section provide very strong evidence for RuCN \(\rightarrow\) Rubpyamide PEnT in D\(_2\)O on the basis of the clear grow-ins of absorbance of bleach and transient signals of the Rubpyamide terminus in TRIR kinetic traces in D\(_2\)O in the \(\Delta t = 0\) to 50 ps time range following laser excitation (Figure 3.5.3.2). When this is considered in conjunction with the approximately synchronous decays measured for the RuCN terminus over the same time period this argument becomes more persuasive.

Evidence for Rubpyamide \(\rightarrow\) RuCN PEnT in CH\(_3\)CN has also been presented (Figure 3.5.3.7) primarily on the basis of the grow-in of absorbance of the RuCN bleach and transient signals over the same \(\Delta t = 0\) to 50 ps timescale. The time constants for PEnT were not precisely determined due to the low instrumental spectral resolution of the PIRATE spectrometer (ca. 8-9 cm\(^{-1}\)) and the limited number of spectra collected in the crucial \(\Delta t = 0\) to 20 ps time range, but approximate values of \(\tau_{\text{PEnT,D}_2\text{O}} \sim \) 5-10 ps and \(\tau_{\text{PEnT,CH}_3\text{CN}} \sim \) 9-15 ps were obtained, indicating that through-
space PEnT occurs rapidly between the two termini and provides evidence of the reversible control of the direction of PEnT in this bimetallic system through manipulation of the solvent composition.

2-Step PEnT Model
The presence of additional fine structure (particularly in the transient region) in the ps-TRIR spectral profile of the RuCN terminus in D₂O (Figure 3.5.3.1) and consequent anomalous appearance of the single-pixel kinetic trace may be rationalised by invoking a photophysical model involving a two-step PEnT process. In D₂O the RuCN terminus possesses the highest ³MLCT energy excited state (Ru₈CN → bpy) of all 3 chromophore types in Ru₈CN-BL₁-Ru₄bpyamide, with the two lower energy chromophore types (Ru₈ → BL₁ and Ru₈ → bpyamide) residing on the Ru₄bpyamide terminus (Figure 3.5.5.1). Thus upon photoexcitation PEnT may feasibly take place from the Ru₈CN chromophore (coloured blue in the schematic) to the Ru₄BL chromophore (coloured green in the schematic) which is intermediate in energy of the three. A further PEnT process is thermodynamically favourable in D₂O from the Ru₄BL₁ chromophore to the Ru₄bpyamide chromophore. The reader’s attention is directed to Figure 3.5.5.1 as the schematic depicts this process in a more elegant and colourful manner! The Figure also illustrates why this process would not be feasible in CH₃CN; namely that the ³MLCT energy of the Ru₄BL₁ chromophore is not intermediate in energy between the ³MLCT excited states of the two termini in this case and is unable to act as a conduit for the PEnT process.

Reinvestigating the spectroscopy and kinetics of the RuCN terminus of RuCN-BL₁-Rubpyamide in D₂O (first presented in Figures 3.5.3.1 and 3.5.3.2, below as Figure 3.5.5.2) and in particular examining the kinetic behaviour of the RuCN transient bands it is quite clear that there is a considerable shift of the transient band to higher wavenumber as Δt increases (Figure 3.5.3.2 inset). By monitoring the single-pixel kinetics at the maxima before and after the band shift (located at 2104 and 2114 cm⁻¹ respectively) it is reasonable to suggest that the two species are interconverting. The 2104 cm⁻¹ feature decays biexponentially with time constants \( \tau_1 = 14 \pm 2 \) ps and \( \tau_2 = \)
292 (±108) ps and the 2114 cm⁻¹ feature exhibits a grow-in over the time range Δt = 0-30 ps followed by a decay with time constant τ₁ = 366 (±60) ps.

Figure 3.5.5.1: Schematic of the chromophores (top) and energy level ordering (bottom) in RuCN-BL'Bpyamide that provides the framework for the 2-step PEnT model.

A shift of 10 cm⁻¹ to higher wavenumber is consistent with excited state electron density being redistributed further away from the RuCN terminus metal centre, as would be expected for a change in the excited chromophore from RuₐBL to Ruₐbpyam (Figure 3.5.5.1). The time constants obtained for these features are in satisfactory agreement with each other. The only obvious alternative explanation for
features on this timescale is vibrational cooling, however the lack of observation of any vibrational cooling of $v(\text{CN})$ features throughout this Thesis renders this suggestion improbable.

![Figure 3.5.5.2: ps-TRIR kinetics of transient bands of RuCN-BL$^1$-Rubpyamide in D$_2$O (main Figure) and ps-TRIR spectra with bands clearly marked (inset).](image)

A simple conformational analysis of RuCN-BL$^1$-Rubpyamide has been performed computationally (§3.5.4) and gave the structures of a number of important conformers of the flexible system, based on the metal centres being either syn or anti to each other and the structure adopting a folded or open geometry (Figure 3.5.4.1). Due to the ease of rotation about the CH$_2$CH$_2$ axis of BL$^1$ it has been postulated that these structures are likely to be members of an ensemble of conformers present in solution, leading to a range of intermetallic distances between $5.41 < r < 12.7$ Å. This provides a plausible explanation for the multiexponential kinetics obtained by luminescence and TRIR spectroscopy as interchromophoric distance and Förster PEnT rates are sensitively interdependent: $k_{\text{PEnT}} \propto r^6$ (§1.1.3). The electrostatic properties of the system are also important to consider; there may be a strong tendency for the complex to ion-pair as the termini are doubly charged with opposite polarity. This may lead to either intra or intermolecular ion pairing, affecting the population of conformers present in solution.
and hence possibly favouring more closed conformers. A study of FTIR lineshapes of RuCN-BL\textsuperscript{1}-Rubpyamide with varying concentration in H\textsubscript{2}O (not shown) was carried out but found no evidence for intermolecular ion-pairing in this system.

The success of the photophysical model in predicting the direction of PEnT in solvents of varying A. N. and the direct evidence for PEnT in both solvents from ps-TRIR spectroscopy is a noteworthy result as, to the best of the author's knowledge at the time of writing, it constitutes the first such case of \textit{real time} direct monitoring of interchromophoric PEnT using ps-TRIR spectroscopy. Furthermore it is certainly the first case of the use of ps-TRIR spectroscopy to probe a solvent-switchable PEnT process in solution.

In order to probe the role that conformation plays in the photophysics of this molecule we have designed and studied a next generation bimetallic complex RuCN-BL\textsuperscript{2}-Rubpyamide in which the freedom of rotation in the bridging ligand is constrained.

3.6. [((4,4'-CON(CH\textsubscript{2}CH\textsubscript{3})\textsubscript{2}-bpy)\textsubscript{2}Ru(bpy-(4,4-CH\textsubscript{2}CH\textsubscript{2})\textsubscript{2}-bpy)Ru(CN)\textsubscript{4}], RuCN-BL\textsuperscript{2}-Rubpyamide

RuCN-BL\textsuperscript{2}-Rubpyamide may be considered a fusion of the two mononuclear complexes RuCN\textsubscript{dmb} (§3.2) and Rubpyamide-BL\textsuperscript{2} (§3.4) and has been designed to exploit properties of these two moieties. As with RuCN-BL\textsuperscript{1}-Rubpyamide, both
termini exhibit lowest-lying $^3\text{MLCT}$ excited states as characterised by absorption, luminescence and TRIR spectroscopy.

The principal difference between this system and RuCN-BL$_1$-Rubpyamide discussed in the previous section is the change in the nature of the bridging ligand, from an open, conformationally flexible structure (BL$^1$) to a closed-cycle system, in which molecular movements are fairly restricted (BL$^2$). This is postulated to lead to a decrease in the range of interchromophoric distances that can be accessed by RuCN-BL$_2$-Rubpyamide, and hence a reduction in the complexity of the kinetic profile of the system, as observed on the nanosecond timescale by luminescence and on the picosecond timescale by TRIR spectroscopy. As the system is also constrained in geometries which allow good spatial overlap between donor and acceptor orbitals, an increase in the efficiency of PEN'T is also expected. The Dexter mechanism is unlikely to operate in this molecule as the saturated CH$_2$CH$_2$ linkers attaching the bridging ligand’s bipyridyls together prevent any direct electronic communication between the two metal centres.

3.6.1. UV/visible Absorption and Luminescence Spectroscopy

UV/visible absorption spectra of RuCN-BL$_2$-Rubpyamide have been measured in CH$_3$CN and D$_2$O (see Appendix §3.9 for spectra) and the peak positions and assignments are presented in Table 3.6.1.1.

<table>
<thead>
<tr>
<th>Solvent</th>
<th>$\lambda_{\text{max}}$ / nm</th>
</tr>
</thead>
<tbody>
<tr>
<td>D$_2$O</td>
<td>&lt;300 359 412 435 470</td>
</tr>
<tr>
<td>CH$_3$CN</td>
<td>&lt;300 355 434 468 507</td>
</tr>
<tr>
<td>Assignment Terminus</td>
<td>IL $^1\text{MC}$ $^1\text{MLCT}$ $^1\text{MLCT}$ $^1\text{MLCT}$</td>
</tr>
<tr>
<td></td>
<td>$\pi$$\pi^*$ RuCN Rubpyamide RuCN</td>
</tr>
</tbody>
</table>

Table 3.6.1.1: Table of peak positions and assignments for UV/visible absorption spectra of RuCN-BL$_2$-Rubpyamide in D$_2$O and CH$_3$CN solutions.
The profiles resemble those obtained for \textbf{RuCN-BL}$_1$-Rubpyamide (§3.5.1) and correspond well to a superposition of the spectra of the model complexes \textbf{RuCN}$_{\text{dmb}}$ and \textbf{Rubpyamide-BL}$_2$, which is consistent with the two containing identical chromophores and little or no electronic communication between the two metal centres in their ground states. As with \textbf{RuCN-BL}$_1$-Rubpyamide, the two relatively solvent-invariant features (<10 nm difference between CH$_3$CN and D$_2$O) can be assigned to Ru → (4.4'-(CONEt$_2$)$_2$bp) $^1\text{MLCT}$ transitions. The less intense longer wavelength feature varies over a range of ca. 85 nm, exhibiting negative solvatochromism as is expected for ruthenium (II) cyanide transitions and is thus assigned to a Ru → ((5.5'-(CH$_3$)$_2$)bp) $^1\text{MLCT}$ transition.

Emission spectra of \textbf{RuCN-BL}$_2$-Rubpyamide in D$_2$O and CH$_3$CN are shown in Figure 3.6.1.1 and 3.6.1.2 respectively. Figure 3.6.1.3 compares emission in the two solvents.

![Emission spectra of RuCN-BL$_2$-Rubpyamide](image)

\textit{Figure 3.6.1.1: Uncorrected emission spectra of RuCN-BL$_2$-Rubpyamide in D$_2$O, with $\lambda_{ex} = 469$ nm (black), $\lambda_{ex} = 424$ nm (red) and $\lambda_{ex} = 410$ nm (green). The spectra were recorded under identical conditions.}
Figure 3.6.1.2: Uncorrected emission spectra of RuCN-BL$^2$-Rubpyamide in CH$_3$CN, with $\lambda_{\text{ex}} = 473$ nm (black), $\lambda_{\text{ex}} = 437$ nm (red) and $\lambda_{\text{ex}} = 400$ nm (green). The spectra were recorded under identical conditions.

As with RuCN-BL$^1$-Rubpyamide, the complex is luminescent giving a broad, featureless spectrum with $\lambda_{\text{max}}$ (D$_2$O) = 657 nm and $\lambda_{\text{max}}$ (CH$_3$CN) = 651 nm. Whilst there is only small variation in the energy of the emissive state, there is a substantial difference in the emission intensity measured in the two solvents; the luminescence is much weaker in CH$_3$CN than in D$_2$O, with the wavelengths and quantum yields presented in Table 3.6.1.2. This is consistent with the emissive states being $^3$MLCT in nature, with the vast majority of the detected luminescence being derived from the Rubpyamide terminus (Rubpyamide-BL$^2$ emits at ca. 650 nm).

These results confirm that the photophysical behaviour of RuCN-BL$^2$-Rubpyamide varies significantly in different solvents. We can rationalise these results in the framework of our model (§3.1) for these complexes. In CH$_3$CN the $^3$MLCT excited state of the Rubpyamide terminus is the higher in energy of the two and thus luminescence is partially quenched by energy transfer to the RuCNdmb terminus, which only emits weakly with barely detectable luminescence. There is a ca. 90%
reduction in quantum yield from that of Rubpyamide-BL\textsuperscript{2} in CH\textsubscript{3}CN. This partial quenching gives us further insight into the efficiency of the PEnT process in this system. In D\textsubscript{2}O the lowest energy $\textsuperscript{3}$MLCT excited state of the Rubpyamide terminus is lower in energy than the $\textsuperscript{3}$MLCT excited state on the RuCN terminus and thus its luminescence remains unquenched at an intensity comparable to that of Rubpyamide-BL\textsuperscript{2}; there is less than 10% difference between the values obtained for RuCN-BL\textsuperscript{2}-Rubpyamide and its amide model complex. The luminescence from the RuCN terminus is extremely difficult to detect in this case as it ostensibly occurs at a wavelength coincident with the intense luminescence from the Rubpyamide terminus.

![Emission Spectra](image)

**Figure 3.6.1.3: Uncorrected emission spectra of RuCN-BL\textsuperscript{2}-Rubpyamide in D\textsubscript{2}O (black) and CH\textsubscript{3}CN (red), with $\lambda_{ex}$ (D\textsubscript{2}O) = 410 nm and $\lambda_{ex}$ (CH\textsubscript{3}CN) = 400 nm. The spectra were recorded under identical conditions.**
A kinetic analysis was performed using 405 nm excitation and registration of emission near the peak maximum (644 nm and 631 nm in D$_2$O and CH$_3$CN respectively). It was found that the traces obtained could not be fitted using a simple exponential function; indeed at least 4 components with time constants varying by orders of magnitude were required to obtain acceptable fits to the experimental data ($\chi^2 < 1.05$). The results are summarised in Table 3.6.1.2 and the attempts to fit the experimental data to between 1 and 4 components in D$_2$O and CH$_3$CN is displayed in Figures 3.6.1.4 and 3.6.1.5 respectively. From these it is clear that at least 4 components are required for satisfactory fits, and in the case of D$_2$O the large deviation of $\chi^2$ from 1 (1.416) and the residual plot demonstrate that a higher number of exponents is required to generate a good fit to the data.

<table>
<thead>
<tr>
<th>Component</th>
<th>Time constant / ns</th>
<th>% Contribution</th>
<th>Time constant / ns</th>
<th>% Contribution</th>
</tr>
</thead>
<tbody>
<tr>
<td>$\tau_1$</td>
<td>5.6 ± 0.6</td>
<td>2.37</td>
<td>3.9 ± 0.21</td>
<td>3.97</td>
</tr>
<tr>
<td>$\tau_2$</td>
<td>50.4 ± 4.1</td>
<td>34.7</td>
<td>12.0 ± 3.2</td>
<td>5.66</td>
</tr>
<tr>
<td>$\tau_3$</td>
<td>100.0 ± 9.3</td>
<td>34.0</td>
<td>100.9 ± 3.2</td>
<td>8.85</td>
</tr>
<tr>
<td>$\tau_4$</td>
<td>483.1 ± 27.9</td>
<td>28.9</td>
<td>742.6 ± 34.9</td>
<td>81.54</td>
</tr>
</tbody>
</table>

Table 3.6.1.2: Table of exponential fit components of luminescence spectra of RuCN-BL$^2$-Rubpyamide in D$_2$O and CH$_3$CN solutions.
Figure 3.6.1.4: Fits to lifetime data (with residuals) for RuCN-BL²-Rubpyamide in D₂O. Top left: 1 exponent; top right: 2 exponents; bottom left: 3 exponents, bottom right: 4 exponents.
Figure 3.6.1.5: Fits to lifetime data (with residuals) for RuCN-BL$^2$-Rubpyamide in CH$_3$CN. Top left: 1 exponent; top right: 2 exponents; bottom left: 3 exponents, bottom right: 4 exponents.
As with RuCN-BL\textsuperscript{1}-Rubpyamide, the kinetic data obtained is clearly polyexponential in form, although it is important to note that $\tau_1$ in both solvents may arise from the IRF. These results suggest that there are still a number of photophysical processes occurring in solution on the nanosecond timescale, despite the system’s decrease in conformational degrees of freedom. TRIR spectroscopy is employed in §3.6.3 to study these processes in more detail.

### 3.6.2. Cyclic Voltammetry

**RuCN-BL\textsuperscript{2}-Rubpyamide** has been studied using cyclic voltammetry in CH\textsubscript{3}CN (see Appendix §3.9 for the voltammogram) and the results are presented in Table 3.6.2.1. The picture is very similar to that of RuCN-BL\textsuperscript{1}-Rubpyamide (§3.5.2), with one reversible oxidation, a large irreversible feature at positive potential (due to oxidation of the amide groups) and three reversible reduction couples arising from the bpy-based ligands.

<table>
<thead>
<tr>
<th>$E_{1/2}/V$ (vs. Fe/Fe\textsuperscript{+})</th>
<th>Reversibility</th>
<th>Assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>+0.96</td>
<td>Reversible</td>
<td>Ru\textsuperscript{II/III}</td>
</tr>
<tr>
<td>($E_p$) +0.70</td>
<td>Irreversible</td>
<td>(4,4'-\text{-(CONEt\textsubscript{2})\textsubscript{2}bpy})\textsuperscript{0/+}</td>
</tr>
<tr>
<td>-1.54</td>
<td>Reversible</td>
<td>(4,4'-\text{-(CONEt\textsubscript{2})\textsubscript{2}bpy})\textsuperscript{0/-}</td>
</tr>
<tr>
<td>-1.75</td>
<td>Reversible</td>
<td>(4,4'-\text{-(CONEt\textsubscript{2})\textsubscript{2}bpy})\textsuperscript{0/-}</td>
</tr>
<tr>
<td>&lt;-2.00</td>
<td>Reversible</td>
<td>(BL\textsuperscript{2})\textsuperscript{0/-}</td>
</tr>
</tbody>
</table>

**Table 3.6.2.1: Table of potentials for redox couples and assignments for the cyclic voltammogram of RuCN-BL\textsuperscript{2}-Rubpyamide in CH\textsubscript{3}CN solution.**

These results are expected from a system containing a Rubpyamide-BL\textsuperscript{1} unit (§3.3.2 3.5.2), as the Rubpyamide redox couples fall within the solvent window and are in very close agreement with those measured for the model complexes. Interestingly the Ru\textsuperscript{II/III} couple of RuCNdmb [$E_{1/2} = -0.03$ V (vs. Fe/Fe\textsuperscript{+})] is expected to be present but

3-76
isn't observed, suggesting that all its reduction processes occur outside the solvent window.

### 3.6.3. FTIR and Time-Resolved Infrared Spectroscopy

Luminescence spectroscopy (§3.6.1) has provided some indirect evidence for through-space PEnT between the two metal centres of RuCN-BL$^2$-Rubpyamide, through the quenching of emission from the Rubpyamide terminus in low A. N. solvent environments. TRIR spectroscopy has been used to monitor this process directly and verify its occurrence. As with RuCN-BL$^1$-Rubpyamide, the system has two sets of IR active 'reporters' - one on either terminus - both can provide important and complementary information and it is crucial to consider them together when analysing the data.

As with RuCN-BL$^1$-Rubpyamide it is important to remember that due to the very similar nature of the chromophores appended to the two metal centres and the broadness of $^1$MLCT absorption peaks, there is often a case of unselective excitation – i.e. either metal centre may be excited by the incident laser light and more than one excited state species may be present in solution simultaneously. This may add another degree of complexity to the study, as bona fide PEnT processes must be disentangled from other primary photophysical processes.

#### Deuterium Oxide Solution

In D$_2$O, SSDA interactions elevate the energy of the lowest-energy RuCN $^3$MLCT state above the corresponding state in the Rubpyamide terminus. Therefore RuCN → Rubpyamide PEnT is expected to take place.

The FTIR spectrum of RuCN-BL$^2$-Rubpyamide in D$_2$O solution in the ν(CN) and amide ν(CO) region has been measured (see Appendix §3.9 for spectra). In the cyanide region there is a peak profile characteristic of the tetracyanoruthenate (II) unit with 3 resolved peak maxima at 2091, 2051 and 2032 cm$^{-1}$, and an additional peak is

3-77
observed at 1995 cm$^{-1}$. In the amide region there is a single peak present at 1615 cm$^{-1}$. These results are comparable with the constituent complexes RuCNdmb and Rubpyamide-BL$^2$ in D$_2$O solution (§3.2.5, §3.4.3).

Picosecond and nanosecond TRIR spectroscopy (with laser excitation at 400 and 355 nm respectively) were performed on RuCN-BL$^2$-Rubpyamide in D$_2$O solution, monitoring both regions in the IR where reporter groups are present.

ps-TRIR Spectroscopy
Figure 3.6.3.1 shows the ps-TRIR spectra obtained for the RuCN and Rubpyamide termini in D$_2$O.

RuCN ("Donor") terminus
The spectral profile obtained following 400 nm excitation contains several overlapped bleaches centred between 2030 and 2090 cm$^{-1}$ and transient bands detected at lower and higher wavenumber to the bleach, centred at 2012 and 2100 cm$^{-1}$. The lower energy bleach is evident at ca. 2000 cm$^{-1}$. There appear to be changes in the profile on the picosecond timescale, particularly in the region where the bleaches overlap with the transient bands. Over the time range of the experiment (1-1000 ps), the majority of the v(CN) bleach and transient signals decay, indicating that there are photophysical processes occurring which facilitate the quenching of the RuCNdmb$^3$MLCT excited state. At $\Delta t = 1$ ns only ca. 10% of the signal immediately after excitation remains, indicating that the majority of RuCNdmb excited states formed in solution have been quenched.
Figure 3.6.3.1: ps-TRIR spectra of RuCN-BL$_2$-Rubpyamide in D$_2$O solution in the v(CN) and amide v(CO) region for $\Delta t = 0$-2000 ps.
Rubpyamide ("Acceptor") terminus

The spectra obtained for the Rubpyamide terminus in D$_2$O resemble the profile observed for the precursor complex Rubpyamide-Blr$^2_1$, with an amide v(CO) bleach centred at 1614 cm$^{-1}$ and a transient at lower wavenumber centred at 1589 cm$^{-1}$. This is consistent with the anticipated $^3$MLCT lowest excited state for this terminus. The spectra show much less development over the timescale of the experiment than for the RuCN terminus, indicating that ultrafast processes are not contributing to excited state quenching to the same extent as in the "donor" terminus. Indeed at $\Delta t = 1$ ns there is still a strong signal remaining, indicating that excited state decay is taking place on longer timescales.

This initial result is consistent with our model, which predicts that the RuCN donor terminus will be quenched by efficient PENt and have a short-lived excited state, whilst the Rubpyamide acceptor terminus will possess a long-lived excited state owing to it being the ultimate excited state being formed, in keeping with results observed for typical systems based on [Ru(bpy)$_3$]$^{2+}$ (§1.4.1.1).

ps-TRIR Kinetics

Figures 3.6.3.2 and 3.6.3.3 show the ps-TRIR kinetics (for bleach and transient where possible) obtained for the RuCN and Rubpyamide termini in different time domains. Figure 3.6.3.2 shows $\Delta t = 0$-50 ps and 3.6.3.3 shows $\Delta t = 0$-1 ns.

$\Delta t = 0$-50 ps

This kinetics recorded during this period do not show any clear evidence for the growth in magnitude of the bleach on the Rubpyamide terminus over the first ca. 30 ps. This could be due to the offsetting of such a process by other processes occurring over a similar timescale (e.g. vibrational cooling) affecting the transient band width and hence affecting the bleach signal due to overlap between the two bands.
Figure 3.6.3.2: ps-TRIR kinetics of RuCN-BL$^2$-Rubpyamide in D$_2$O solution for the RuCN (transient at 2103 cm$^{-1}$, top left and bleach at 2051 cm$^{-1}$, bottom left) and Rubpyamide (transient at 1592 cm$^{-1}$, top right and bleach at 1613 cm$^{-1}$, bottom right) termini for $\Delta t = 0$-50 ps.
Figure 3.6.3.3: ps-TRIR kinetics of RuCN-BL$^2$-Rubpyamide in D$_2$O solution for the RuCN (transient at 2103 cm$^{-1}$, top left and bleach at 2051 cm$^{-1}$, bottom left) and Rubpyamide (transient at 1592 cm$^{-1}$, top right and bleach at 1613 cm$^{-1}$, bottom right) termini for $\Delta t = 0$-1 ns.
Depletion in the magnitude of both the bleach and the transient on the RuCN terminus does occur over this same timescale, indicating that a quenching process, probably PEnT, is operating in this time domain. The signal intensity is such that precise time constants cannot be analysed meaningfully, however they are all within the region of 8-20 ps, which would correspond to a time constant of 12 (±8) ps for efficient and rapid PEnT in this supramolecular system.

\[ \Delta t = 0-1 \text{ ns} \]

On this timescale the processes described above are observed following excitation, and these are followed by signal decay on both termini, indicating that the excited state decay by routes other than PEnT has become the major process occurring. As was observed for RuCN-BL\textsuperscript{1}-Rubpyamide, these processes appear to be multi-exponent. This may be due to the offsetting of excited state decay processes with PEnT-related grow-in of signal; or the opposite situation where there are multiple competing decay processes occurring at varying rates. The time constant for the major decay component in each case is in the region of 400-700 ps.

**ns-TRIR Spectroscopy and Kinetics**

In contrast to the situation for the RuCN-BL\textsuperscript{1}-Rubpyamide, both termini have sufficiently long-lived excited states in D\textsubscript{2}O to allow monitoring on the nanosecond timescale. Figures 3.6.3.4 and 3.6.3.5 show the spectroscopy and kinetics measured at \( \Delta t = 1-300 \text{ ns} \).

The ns-TRIR spectroscopy of the RuCN terminus following excitation at 355 nm is analogous to the ps-TRIR spectroscopy shown above, with several overlapping bleaches centred at \( \text{ca. } 2050 \text{ cm}^{-1} \) and transient features centred at 2105 and 2020 cm\textsuperscript{-1}. The ns-TRIR spectrum of the Rubpyamide terminus contains a bleach centred at 1615 cm\textsuperscript{-1} and a transient band centred at 1590 cm\textsuperscript{-1}. The signals detected were weak, implying that much of the excited state had decayed on the picosecond timescale, although the signal intensities over the two experiments are unfortunately not directly
comparable due to different laser excitation wavelengths, data collection times and sample concentrations.

Analysis of the kinetics at the bleach and transient peak maxima give multiexponent fits, with good fits obtained using 2 exponents. The RuCN (donor) terminus has time constants of 3.2 (±1.8) and 19 (±9.2) ns, and the Rubpyamide (acceptor) terminus has time constants of 8.4 (±3.7), 45.1 (±9.9) and 138 (±50) ns. This is consistent with the luminescence data presented in §3.6.1 and confirms that there are multiple photophysical processes occurring in this system over the timescales investigated.

**Acetonitrile Solution**

In CH₃CN, SSDA interactions stabilise the energy of the RuCN ³MLCT state so that it is lower than the corresponding state on the Rubpyamide terminus. Therefore Rubpyamide → RuCN PEnT is expected to take place.

The FTIR spectrum of RuCN-BL²-Rubpyamide has been measured in CH₃CN solution in the ν(CN) and amide ν(CO) region (see Appendix §3.9 for spectra). 3 components to the peak are resolved at 2086, 2069 and 2058 cm⁻¹. Interestingly there appears to be an additional band present at 1989 cm⁻¹. In the amide region there is a single peak present at 1637 cm⁻¹. These results are broadly comparable with the constituent complexes RuCNdmb and Rubpyamide-BL² in CH₃CN solution (§3.2.5, §3.4.3) but there are significant differences to the corresponding RuCN-BL¹-Rubpyamide FTIR spectrum. The relative intensities of the components are ostensibly reversed in the two complexes, giving the spectra a ‘mirror image’ type relationship. Secondly the additional peak at lower wavenumber would be expected to bleach and give rise to transient peaks in the TRIR experiments if it is a ν(CN) band.
Figure 3.6.3.4: ns-TRIR spectra of RuCN-BL₂-Rubpyamide in D₂O solution for both termini for Δt = 0-300 ns.
Figure 3.6.3.5: ns-TRIR kinetics of RuCN-BL\(^2\)-Rubpyamide in D\(_2\)O solution for RuCN (transient at 2102 cm\(^{-1}\), top left and bleach at 2052 cm\(^{-1}\), bottom left) and Rubpyamide (transient at 1585 cm\(^{-1}\), top right and bleach at 1618 cm\(^{-1}\), bottom right) termini for \(\Delta t = 0-300\) ns.
Picosecond and nanosecond TRIR spectroscopy (with laser excitation at 400 and 355 nm respectively) were performed on RuCN-BL$_2$-Rubpyamide in CH$_3$CN solution, monitoring both regions in the infrared spectral region where reporter groups are present.

**ps-TRIR Spectroscopy**

Figure 3.6.3.6 shows the ps-TRIR spectra obtained for the RuCN and Rubpyamide termini in CH$_3$CN solution.

**RuCN ("Acceptor") terminus**

The spectral profile obtained following 400 nm excitation in CH$_3$CN consists of several overlapped bleaches centred between 1975 and 2100 cm$^{-1}$ and transient bands with maxima at 2016 and 2112 cm$^{-1}$. These originate from the 1989 and 2058 cm$^{-1}$ features respectively. No major changes in the spectral profile are apparent on the picosecond timescale. Over the time range of the experiment (1-1000 ps), the majority of the ν(CN) bleach and transient signals decay, although to a lesser extent than was observed in D$_2$O. At Δt = 1 ns there is still a significant residual signal remaining, indicating that the quenching of RuCN-dmb excited states formed in CH$_3$CN is much less efficient than in D$_2$O.

**Rubpyamide ("Donor") terminus**

The spectra obtained for the Rubpyamide terminus in CH$_3$CN resemble the profile observed for the precursor complex Rubpyamide-BL$_2$, with an amide ν(CO) bleach centred at 1637 cm$^{-1}$ and a transient to lower wavenumber centred at 1616 cm$^{-1}$. This is consistent with a $^3$MLCT excited state for this terminus. In contrast to the observation in D$_2$O, there is a much greater extent of transient decay and parent recovery on the picosecond timescale, such that over 90 % of the integrated absorbance of the bleach has decayed by Δt = 1 ns.

This result is consistent with our model, which predicts that the Rubpyamide donor terminus will be quenched by efficient PEnT and have a short-lived excited state,
whilst the RuCN acceptor terminus will possess a long-lived excited state owing to it being the ultimate excited state being formed. The slight differences in the proportion of signal intensity remaining at $\Delta t = 1$ ns for the two termini between $D_2O$ and CH$_3$CN can be understood by taking into account the very different natural excited state lifetimes of the two units (as measured in the model complexes RuCN$_{dmb}$ and Rubpyamide-BL$_2$, §3.2.5, §3.4.3) and indeed the very significant effect that solvent composition has on the excited state properties of RuCN$_{dmb}$ in particular.

### ps-TRIR Kinetics

Figures 3.6.3.7 and 3.6.3.8 show the ps-TRIR kinetics (for bleach and transient where possible) obtained for the RuCN and Rubpyamide termini in different time domains in CH$_3$CN solution. Figures 3.6.3.7 shows $\Delta t = 0-50$ ps and 3.6.3.8 shows $\Delta t = 0-1$ ns.

#### $\Delta t = 0-50$ ps

This kinetics recorded during this period are similar to those presented for RuCN-BL$_1$-Rubpyamide, as there is some evidence for the growth in magnitude of the bleach and transient on the RuCN terminus over the first ca. 30 ps, but unfortunately no data of comparable clarity for the Rubpyamide terminus were obtained. For the RuCN terminus, which as the acceptor provides the tell-tale grow-in of absorbance on both bleach and transient bands, the time constants obtained cannot be trusted precisely (owing to poor data) but they are all within the region of 0.5-3 ps. This would correspond to a time constant of ~2 (±1) ps for efficient and rapid PEnT in this supramolecular system. Analysis of the donor terminus data provides few conclusions as neither a clear grow-in or decay is observed over the $\Delta t = 0-50$ ps timescale. Possible reasons for this include offsetting of decay processes in the Rubpyamide terminus by peak sharpening of the amide bands due to vibrational cooling processes.
Figure 3.6.3.6: ps-TRIR spectra of RuCN-BL₄²⁻-Rubpyamide in CH₃CN solution in the ν(CN) and amide ν(CO) region for Δt = 0-2000 ps.
\[ \Delta t = 0.1 \text{ ns} \]

As was observed in D_{2}O solution, during this time domain the processes described above are followed by signal decay on both termini, indicating that the excited state decay by mechanisms other than PEnT becomes the major process occurring. These processes appear to be multi-exponential, with fits employing 1 or 2 components most closely representing the data. This may be due to the offsetting of excited state decay processes with PEnT related grow-in of signal – or the converse situation where there are multiple competing decay processes occurring at varying rates. There appears to be little consistency in the time constants determined for the fits to the experimental data, with time constants varying between 85 and 250 ps for the major decay component.

**ns-TRIR Spectroscopy and Kinetics**

In CH_{3}CN our photophysical model predicts that the RuCN terminus will have a long-lived excited state which can be monitored on the nanosecond timescale. Figures 3.6.3.9 and 3.6.3.10 show the spectroscopy and kinetics measured at \( \Delta t = 1-500 \text{ ns} \). ns-TRIR spectra of the Rubpyamide terminus contained no excited state signatures in CH_{3}CN solution, conceivably due to its short excited state lifetime.

The ns-TRIR spectra of the RuCN terminus following excitation at 355 nm are similar to the ps-TRIR spectra shown above, with several overlapping bleaches centred between 2050 and 2085 cm\(^{-1}\) and transient bands at lower and higher wavenumber centred at 2023 and 2095 cm\(^{-1}\). The signals detected were weak, implying that most of the excited state had decayed on the picosecond timescale, although the signal intensities over the two experiments are unfortunately not directly comparable due to different laser excitation wavelengths, data collection times and sample concentrations. Interestingly the ns-TRIR spectra suggest that there is a strong transient signal to lower wavenumber than the main parent bleach band (centred at 2019 cm\(^{-1}\)), however this may be an artefact arising from the poor signal-to-noise ratios obtained as the excited state decays.
Figure 3.6.3.7: ps-TRIR kinetics of RuCN-BL$_2$-Rubpyamide in CH$_3$CN solution for the RuCN (transient at 2107 cm$^{-1}$, top left and bleach at 2056 cm$^{-1}$, bottom left) and Rubpyamide (transient at 1614 cm$^{-1}$, top right and bleach at 1638 cm$^{-1}$, bottom right) termini for $\Delta t = 0$-50 ps.
Figure 3.6.3.8: ps-TRIR kinetics of RuCN-BL₂-Rubpyamide in CH₃CN solution for the RuCN (transient at 2107 cm⁻¹, top left and bleach at 2056 cm⁻¹, bottom left) and Rubpyamide (transient at 1614 cm⁻¹, top right and bleach at 1638 cm⁻¹, bottom right) termini for Δt = 0-1 ns.
Analysis of the kinetics at the bleach and transient peak maxima give biexponential fits, with components having time constants of 4.2 (±0.9) and 110 (±41) ns. This is unfortunately incomparable with the corresponding luminescence experiment as the weak RuCN luminescence was undetectable alongside the much more intense emission from the Rubpyamide terminus. This confirms that there are multiple photophysical processes occurring in this system over the timescales investigated.

As discussed in §3.5.3, the processes above have been confirmed as intramolecular through experiments on solution mixtures of the model complexes RuCNdmb and Rubpyamide-BL\textsuperscript{1}.

### 3.6.4. Computational Studies and Conformational Analysis

RuCN-BL\textsuperscript{2}-Rubpyamide has been designed to remove the large amount of conformational freedom that RuCN-BL\textsuperscript{1}-Rubpyamide possesses in solution. Molecular mechanics calculations have been performed using the SPARTAN02 Computational Suite to gain insight into the effect that this change in ligand design has on the parameters that govern the rate and efficiency of PEnT.

Whilst the BL\textsuperscript{2} ligand has a limited amount of conformational flexibility, two minimum energy conformers have been found (Figure 3.6.4.1). It should be noted that the thermodynamic barriers to interconversion are likely to be inaccessible in room temperature solution, and the conformational ratio is likely to be decided during the synthesis of the complex, either during cyclisation of the ligand upon complexation of metal centres. Conformer (i) is a ‘roofed’ structure, in which the BL\textsuperscript{2} ligand resembles an open book, with the alkane linkers forming a plane (the ‘spine’) and the bpy ligands forming the ‘front and back covers’. Conformer (ii) is an ‘open’ structure, where the BL\textsuperscript{2} ligand is unfolded and the two metal centres point outwards.
Figures 3.6.3.9 and 3.6.3.10: ns-TRIR spectra (left) and kinetics (transient at 2016 cm\(^{-1}\), top left and bleach at 2054 cm\(^{-1}\), bottom left) of RuCN-BL\(^2\)-Rubpyamide in CH\(_3\)CN solution for the RuCN terminus for \(\Delta t = 0\)-300 ns.
The calculated intermetallic distances in these conformers are 8.8 Å in conformer (i) and 11.9 Å in conformer (ii). As the there is expected to be no interconversion between the two, and limited conformational freedom for each case, a bimodal distribution of intermetallic distances would be expected, centred around 8.8 and 11.9 Å.

As with RuCN-BL\(^1\)-Rubpyamide there is a significant electrostatic attraction between the two metal centres. This is of less importance in BL\(^2\) bridged systems as the lack of conformational freedom reduces the likelihood of intramolecular ion-pairing.

![Figure 3.6.4.1: Two important conformers of RuCN-BL\(^2\)-Rubpyamide calculated in the gas phase using SPARTAN's molecular mechanics package. i) Represents a ‘roofed’ conformer, ii) is an open structure with both metal centres pointing outwards.](image)
3.6.5. RuCN-BL\textsuperscript{2}-Rubpyamide: Closing Points

The bimetallic complex \textbf{RuCN-BL\textsuperscript{2}-Rubpyamide} has been studied using UV/visible absorption and luminescence spectroscopy, cyclic voltammetry and TRIR spectroscopy in D\textsubscript{2}O and CH\textsubscript{3}CN solutions.

UV/visible spectroscopy (§3.6.1) and cyclic voltammetry (§3.6.2) of \textbf{RuCN-BL\textsuperscript{2}-Rubpyamide} have shown that the ground state properties of the system are essentially a superposition of the constituent monometallic fragments most closely represented by the model complexes \textbf{RuCNdmb} and \textbf{Rubpyamide-BL\textsuperscript{2}}. In an analogous manner to \textbf{RuCN-BL\textsuperscript{1}-Rubpyamide}, this supports our supposition that the saturated spacer connecting the two bpy ligands in the bridging ligand BL\textsuperscript{2} prevents electronic communication between the two termini. UV/vis absorption spectroscopy has observed a measurable difference in the absorption profiles of \textbf{RuCN-BL\textsuperscript{2}-Rubpyamide} in CH\textsubscript{3}CN and D\textsubscript{2}O which is more subtle than observed in the model complex \textbf{RuCNdmb}. This is presumably due to the presence of one terminus exhibiting strong negative solvatochromism (RuCN) and one which is essentially solvent-independent (Rubpyamide). These results are consistent with those obtained by Ward and co-workers in the study of \textbf{RuCN-BL-Rubpy} (§3.1) and those presented earlier in this Chapter for \textbf{RuCN-BL\textsuperscript{1}-Rubpyamide} (§3.5.5).

Luminescence spectroscopy of \textbf{RuCN-BL\textsuperscript{2}-Rubpyamide} has shown that the spectral profile is dominated by emission arising from the \textbf{Rubpyamide} terminus. The considerable solvent-dependence of the emission quantum yield and multi-exponent nature of the lifetime has been studied with experiments in CH\textsubscript{3}CN and D\textsubscript{2}O. Solvent-based quenching of emission ($\phi_{(D\textsubscript{2}O)} - \phi_{(CH\textsubscript{3}CN)}$) is greater than in \textbf{RuCN-BL\textsuperscript{1}-Rubpyamide}, consistent with \textbf{Rubpyamide $\rightarrow$ RuCN PEnT} being thermodynamically favourable in low A. N. solvents such as CH\textsubscript{3}CN, and ostensibly that the PEnT process may be more efficient in this system. The presence of multiple time constants has not been affected despite the introduction of the confirmationally
restricted bridging ligand BL₂ although there are still a number of possible conformers which may be present in solution (§3.7.4, vide infra).

TRIR spectroscopy has been used to directly monitor the processes occurring following photoexcitation of RuCN-BL₂-Rubpyamide on the picosecond and nanosecond timescales. As with RuCN-BL₁-Rubpyamide the ps and ns-TRIR data exhibits polyexponential time dependence, supporting the results from emission spectroscopy that the substantial decrease in conformational freedom has apparently not reduced the number of photophysical processes occurring. The Chapter Discussion (§3.9) contains a detailed treatment on alternative explanations for the complexity of the kinetics observed in this Chapter.

In agreement with the model presented in §3.6.3, the results presented in the same section provide some evidence for RuCN → Rubpyamide PEnT in D₂O in the Δt = 0-50 ps time range (Figure 3.6.3.2), but not as definitively as in RuCN-BL₁-Rubpyamide as no clear grow-in of absorbance is observed in the spectra of the Rubpyamide (PEnT acceptor) terminus. Evidence for Rubpyamide → RuCN PEnT in CH₃CN has also been presented (Figure 3.6.3.7) primarily on the basis of the grow-in of absorbance of the RuCN bleach and transient signals over the same Δt = 0-50 ps timescale. The time constants for PEnT were not precisely determined due to the low instrumental spectral resolution of the PIRATE spectrometer (ca. 8-9 cm⁻¹) and the limited number of spectra collected in the crucial Δt = 0-20 ps time range, but an approximate value of τPEnT,CH₃CN ∼ 0.5-3 ps was obtained. It was not possible to obtain an approximate time constant for τPEnT,D₂O; ostensibly this process may be complete within the instrument response of the PIRATE apparatus (ca. 1 ps, see Chapter 6 for more details on the PIRATE experimental apparatus). Considering that for RuCN-BL₁-Rubpyamide τPEnT,D₂O ∼ 0.5 x τPEnT,CH₃CN this explanation seems plausible. These values represent a considerably higher PEnT rates than that determined for RuCN-BL₁-Rubpyamide, presumably due to the rigid BL₂ ligand pre-organising the two chromophores for PEnT with good orbital overlap. Time-resolved spectroscopic
experiments in the femtosecond time domain may provide more insight into the
detailed nature and timescales of these rapid processes.

A simple conformational analysis of \textbf{RuCN-BL}^2-\textbf{Rubpyamide} has been performed
computationally (§3.6.4) and gave the structures of two important conformers of the
system, based on the bridging ligand adopting either a ‘roofed’ or ‘open’ geometry
(Figure 3.6.4.1). The presence of two species in solution may provide a partial
explanation for the multiexponential kinetics obtained by luminescence and TRIR
spectroscopy as interchromophoric distance and and predicted Forster PEnT rates
exhibit a sensitive interdependence (§1.3), but the substantial decrease in degrees of
freedom as a result of replacing BL$^1$ with BL$^2$ has not been accompanied by a
reduction in the complexity of the kinetics of \textbf{RuCN-BL}^2-\textbf{Rubpyamide}. The
electrostatic properties of the system are also important to consider; as was the case
with \textbf{RuCN-BL}^1-\textbf{Rubpyamide} there may be a strong tendency to ion-pair as the
termini are doubly charged with opposite polarity. In the case of \textbf{RuCN-BL}^2-
\textbf{Rubpyamide} the rigidity of BL$^2$ means that intramolecular ion-pairing is unlikely,
therefore intermolecular ion-pairing might be a plausible consequence of this effect.

The success of the photophysical model in predicting the direction of PEnT in
solvents of varying A. N. and the evidence for PEnT in both solvents from ps-TRIR
spectroscopy has been presented and comparisons have been made between the two
solvent-switchable PEnT bimetallic systems in §3.5 and §3.6. In order to study PEnT
processes involving the ligand system BL$^1$ in more detail we have designed and
studied a fixed-energy bimetallic complex \textbf{ReCO-BL}^1-\textbf{Rubpyamide} which is
comprised of a \textbf{Rubpyamide} terminus and a [ReCl(CO)$_3$] unit.
ReCO-BL\textsuperscript{1}-Rubpyamide is a combination of the two mononuclear complexes fac-[ReCl(CO)\textsubscript{3} (bpy)] (§1.4.1.2) and Rubpyamide-BL\textsuperscript{1} (§3.3) and has been designed to exploit properties of these two moieties. Both termini exhibit lowest-lying MLCT excited states as characterised by absorption, luminescence and TRIR spectroscopy. The key difference between this system and the related systems containing the RuCN moiety is that neither chromophore in this system is solvatochromic, enabling us to study the nuances of systems bridged by the BL\textsuperscript{1} ligand with one of the principal variables removed. Furthermore, the electrostatic properties of this system are very different. Whilst there is an electrostatic attraction between the two chromophores in RuCN-BL\textsuperscript{1}-Rubpyamide (as one is doubly cationic and the other is doubly anionic) here there will be no such effect (as one is neutral whilst the other is doubly cationic). This is likely to affect the conformational dynamics of the system in solution, possibly reducing the energetic favourability of the "ion-paired" arrangement which brings the chromophores close together for optimal through-space PEnT (§3.7.4)

Due to the absence of the RuCN unit, the direction and gradient of PEnT in this system should not be affected by solvent. The emission wavelengths of the chromophores [ReCl(CO)\textsubscript{3} (bpy)] (hereafter ReCO) and Rubpyamide are 612 and
641 nm respectively (§1.4.1.2), giving an energy gradient of 740 cm⁻¹ (or 8.9 kJ mol⁻¹) for ReCO → Rubpyamide PEnT.

The ν(CO) bands of the amide groups and the ν(CO) bands of the carbonyl ligands provide complementary vibrational handles in different regions of the IR spectrum, enabling direct monitoring of the evolution of the excited state structure on both metal centres. As in RuCN-BL¹-Rubpyamide, the BL¹ bridging ligand allows the internuclear distance between the metal centres to be modulated. Efficient energy transfer via the Förster mechanism requires that the centres are spatially proximal and that there is little overlap in the emission energy of the donor and absorption energy of the acceptor chromophores.

### 3.7.1. UV/visible Absorption and Luminescence Spectroscopy

The UV/visible absorption spectrum of ReCO-BL¹-Rubpyamide has been measured in CH₃CN (see Appendix §3.9 for spectra) and the peak positions and assignments are presented in Table 3.7.1.1.

<table>
<thead>
<tr>
<th>λ max / nm</th>
<th>&lt; 300</th>
<th>327</th>
<th>350</th>
<th>397</th>
<th>434</th>
<th>468</th>
</tr>
</thead>
<tbody>
<tr>
<td>Assignment</td>
<td>IL</td>
<td>¹MC</td>
<td>¹MC</td>
<td>¹MLCT</td>
<td>¹MLCT</td>
<td>¹MLCT</td>
</tr>
<tr>
<td>Terminus</td>
<td>π-π*</td>
<td>Re</td>
<td>Ru</td>
<td>ReCO</td>
<td>Rubpyamide</td>
<td>Rubpyamide</td>
</tr>
</tbody>
</table>

*Table 3.7.1.1: Table of peak positions and assignments for the UV/visible absorption spectrum of ReCO-BL¹-Rubpyamide in CH₃CN solution.*

The profile is composed of features found in the spectra of the model complexes [ReCl(CO)₃(bpy)] (§1.4.1.2) and Rubpyamide-BL¹ (§3.3.1), which is consistent with little or no electronic communication between the two metal centres in the ground state.

As with RuCN-BL¹-Rubpyamide, the two relatively solvent-invariant features (<10 nm difference between CH₃CN and D₂O) can be assigned to Ru → (4,4¹-
(CONEt$_2$)$_2$bpy) $^1$MLCT transitions. The weaker shorter-wavelength feature can be assigned to a Re $\rightarrow$ bpy $^1$MLCT transition by comparison with the spectrum of [ReCl(CO)$_3$(bpy)] in CH$_3$CN (§1.4.1.2).

Emission spectra of ReCO-BL$^1$-Rubpyamide recorded in CH$_3$CN are shown in Figure 3.7.1.1. As with RuCN-BL$^1$-Rubpyamide, the complex is luminescent giving a broad, featureless spectrum with $\lambda_{\text{max}} = 645$ nm. As with the RuCN-BL$^3$-Rubpyamide systems, this is consistent with the emissive states being $^3$MLCT in nature, with the vast majority of the detected luminescence being derived from the Rubpyamide terminus (Rubpyamide-BL$^2$ emits at ca. 650 nm). The results are presented in Table 3.7.1.1.

![Figure 3.7.1.1: Uncorrected emission spectra of ReCO-BL$^1$-Rubpyamide in CH$_3$CN solution, with $\lambda_{\text{ex}} = 475$ nm (black), $\lambda_{\text{ex}} = 435$ nm (red) and $\lambda_{\text{ex}} = 396$ nm (green). The spectra were recorded under identical conditions.](image)

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A kinetic analysis was performed on the luminescence of ReCO-BL\textsuperscript{1}-Rubpyamide in CH\textsubscript{3}CN using 405 nm excitation and registration of emission at 644 nm with an Edinburgh Instruments FLS920 spectrometer (with time resolution of ca. 5-7 ns). As with the RuCN-BL\textsuperscript{x}-Rubpyamide systems, it was found that the traces obtained could not be fitted using a simple exponential function; 3 components with time constants varying by orders of magnitude were required to obtain acceptable fits to the experimental data ($\chi^2 < 1.07$). The results are summarised in Table 3.7.1.2 and the attempts to fit the experimental data to between 1 and 3 components in CH\textsubscript{3}CN is displayed in Figure 3.7.1.2. It is conceivable that the detection of 3 components (compared with 4/5 components in RuCN-BL\textsuperscript{x}-Rubpyamide) is because an IRF component is not observed because of the comparatively intense luminescence in this case.

<table>
<thead>
<tr>
<th>Component</th>
<th>Time constant / ns</th>
<th>% Contribution</th>
</tr>
</thead>
<tbody>
<tr>
<td>CH\textsubscript{3}CN ($\chi^2 = 1.065$)</td>
<td></td>
<td></td>
</tr>
<tr>
<td>$\tau_1$</td>
<td>$27 \pm 1.2$</td>
<td>19.9</td>
</tr>
<tr>
<td>$\tau_2$</td>
<td>$428 \pm 21$</td>
<td>25.2</td>
</tr>
<tr>
<td>$\tau_3$</td>
<td>$1150 \pm 83$</td>
<td>54.9</td>
</tr>
</tbody>
</table>

Table 3.7.1.2: Table of exponential fit components of luminescence spectra of ReCO-BL\textsuperscript{1}-Rubpyamide in CH\textsubscript{3}CN solution.
Further experiments were carried out to study the kinetics of **ReCO-BL\(^1\)-Rubpyamide** in CH\(_3\)CN solution using 405 nm excitation and registration of emission at 644 nm with a Edinburgh Instruments Mini-Tau spectrometer (with maximum time resolution of ca. 250 ps). This allows monitoring of the 0-5 ns time region (within the IRF of the FLS920 apparatus) during which PEnT is most likely to take place, by comparison with the **RuCN-BL\(^3\)-Rubpyamide** systems presented elsewhere in this Chapter. The trace is shown in Figure 3.7.1.3 and clearly demonstrates that the emission (originating from the **Rubpyamide** terminus as before) can be seen to grow in over the first few nanoseconds following the laser excitation pulse before stabilising and decaying on longer timescales (as observed in the longer timescale experiments described above). Reasonable fits to the rise and decay parts of the trace may be obtained with time constants \(\tau_{\text{rise}} = 0.32 \pm 0.04\) ns and \(\tau_{\text{decay}} = 0.15 \pm 0.05\) ns. \(\tau_{\text{rise}}\) may correspond to the rate of **ReCO \(\rightarrow\) Rubpyamide** PEnT but further evidence is required to assess the validity of this hypothesis.

The results show that a much smaller degree of luminescence quenching exists in **ReCO-BL\(^1\)-Rubpyamide** than in the **RuCN-BL\(^3\)-Rubpyamide** systems discussed earlier, even after taking the effect of solvent-cyanide SSDA interactions on luminescence intensity in CH\(_3\)CN (§3.5.1, 3.6.1) into account. This could be due to efficient PEnT or because of a larger difference in the energy levels of the lowest excited states on the two chromophores reducing the likelihood of thermal deactivation of luminescent states. TRIR spectroscopy is employed to probe this further (§3.7.3). TCSPC has ostensibly monitored the PEnT process, by virtue of the use of apparatus with improved time resolution and the ostensible retardation of the energy transfer process in this heterobimetallic dyad in comparison with the homobimetallic **RuCN-BL\(^3\)-Rubpyamide** dyads previously discussed.
Figure 3.7.1.2: Fits to lifetime data (with residuals) for ReCO-BL¹-Rubpyamide in CH₃CN. Top left: 1 exponent; top right: 2 exponents; bottom: 3 exponents.
Figure 3.7.1.3: Kinetic trace of ReCO-BL₁-Rubpyamide in CH₃CN between 0 and 6 ns.

3.7.2. Cyclic Voltammetry

ReCO-BL₁-Rubpyamide has been studied using cyclic voltammetry in CH₃CN (see Appendix §3.9 for the voltammogram) and the results are presented in Table 3.7.2.1. There are similarities with the two RuCN-BL₁-Rubpyamide systems (§3.5.2, §3.6.2), with a large irreversible feature at positive potential (due to oxidation of the amide groups) and three reversible reduction couples arising from the bpy-based ligands. One important difference is that the couple at $E_p = +0.96$ V is irreversible in this case, suggesting that it is comprised of both Ru and Re oxidation processes. Furthermore, square wave voltammetry results (not shown) suggest that there may be two oxidations overlapping at similar potentials. A plausible explanation for this is that both the ruthenium (II) and rhenium (I) metal centres are being oxidised at similar potentials.
<table>
<thead>
<tr>
<th>$E_{\text{v}}/V$ (vs. Fc/Fc$^+$)</th>
<th>Reversibility</th>
<th>Assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>$(E_p) +0.96$</td>
<td>Irreversible</td>
<td>Ru$^{II/III}$ and Re$^{III}$</td>
</tr>
<tr>
<td>$(E_p) +0.70$</td>
<td>Irreversible</td>
<td>$(4,4'-(CONEt_2)bppy)^{0/+}$</td>
</tr>
<tr>
<td>-1.54</td>
<td>Reversible</td>
<td>$(4,4'-(CONEt_2)bppy)^{0/-}$</td>
</tr>
<tr>
<td>-1.75</td>
<td>Reversible</td>
<td>$(4,4'-(CONEt_2)bppy)^{0/-}$</td>
</tr>
<tr>
<td>&lt;=-2.00</td>
<td>Reversible</td>
<td>$(BL^1)^{0/-}$</td>
</tr>
</tbody>
</table>

*Table 3.7.2.1: Table of potentials for redox couples and assignments for the cyclic voltammogram of ReCO-BL$^1$-Rubpyamide in CH$_3$CN solution.*

These results are expected from a bimetallic complex containing a Rubpyamide-BL$^x$ unit with metal centres exhibiting no ground-state interactions (§3.5.2 3.6.2), as the Rubpyamide redox couples fall within the solvent window and are in very close agreement with those measured for the model complex. No reduction processes originating from the ReCO terminus have been observed.

### 3.7.3. FTIR and Time-Resolved Infrared Spectroscopy

As with the other bichromophoric complexes in this Chapter, it is important to remember that due to the similar nature of the MLCT chromophores appended to the two metal centres and the broadness of $^1$MLCT absorption peaks, there is often a case of unselective excitation. Either metal centre may be excited by the incident laser light and more than one excited state species may be present in solution following excitation. This adds another degree of complexity to the study, as bona fide PEnT processes must be disentangled from other primary photophysical processes.

Due to the insolubility of ReCO-BL$^1$-Rubpyamide in aqueous media it has not been possible to characterise this system in D$_2$O. However this is less crucial than in the supramolecular systems containing the RuCN moiety, as the lack of solvatochromism would be expected to lead to similar results being observed in different solvents.
Acetonitrile Solution

Using measurements of the excited state energies of the two chromophores (from luminescence, §1.4.2.2 and §3.2.1), in CH$_3$CN ReCO $\rightarrow$ Rubpyamide PEnT is expected to take place, and from emission kinetics is postulated to occur in the time range 0-2 ns (§3.7.1).

The FTIR spectrum of ReCO-BL$^1$-Rubpyamide has been measured in CH$_3$CN solution (see Appendix §3.9 for spectra). In the carbonyl region there is a peak profile characteristic of the rhenium (I) tricarbonyl unit with 3 peak maxima at 2023, 1916 and 1897 cm$^{-1}$. In the amide region there is a single peak present at 1637 cm$^{-1}$. These results are closely comparable with the constituent complexes [ReCl(CO)$_3$(bpy)] and Rubpyamide-BL$^1$ in CH$_3$CN solution (§1.4.1.2, §3.3.5).

ReCO-BL$^1$-Rubpyamide was studied using ps-TRIR and ns-TRIR spectroscopy (with laser excitation at 400 and 355 nm respectively) in CH$_3$CN solution, monitoring both regions in the IR where reporter groups are present.

ps-TRIR Spectroscopy

Figure 3.7.3.1 shows the ps-TRIR spectra obtained for the ReCO and Rubpyamide termini in CH$_3$CN solution.

ReCO ("Donor") terminus

The spectral profile obtained following 400 nm excitation resembles that of [ReCl(CO)$_3$(bpy)] in CH$_3$CN, with three bleaches at 2023, 1916 and 1897 cm$^{-1}$ and three transient bands shifted to higher wavenumber at 2058, 1996 and 1959 cm$^{-1}$. There appear to be changes in the profile on the picosecond timescale, particularly in the region where the high energy bleach is overlapping with the transient bands. These may be partially due to vibrational cooling of initially formed $\nu>1$ states, but the kinetic data obtained does not show any definitive evidence for this in the appropriate timescale ($\tau\sim$10 ps, vide infra). Over the time range of the experiment (1-
there is some loss of signal intensity (ca. 30%) of the ReCO ν(CO) bleach and transient signals. At Δt = 1 ns there is still a significant residual signal remaining, indicating that the quenching of ReCO excited states formed in CH3CN is not complete during the picosecond time domain.

**Rubpyamide ("Acceptor") terminus**
The spectra obtained for the Rubpyamide terminus in CH3CN resemble the profile observed for the precursor complex Rubpyamide-BL1 (§4.3.5), with an amide ν(CO) bleach centred at 1635 cm⁻¹ and a transient to lower wavenumber centred at 1625 cm⁻¹. This is consistent with a ³MLCT excited state for this terminus. Very little change in the spectral profile occurs over the duration of the picosecond time window, such that over 90% of the initial signal intensity remains at Δt = 1 ns. This suggests that little quenching of the Rubpyamide excited state is occurring over this timescale. Heavy overlap between the bleach and transient bands may lead to the masking of processes, particularly when two such processes take place over a similar timescale, *e.g.* vibrational cooling and fast PEnT.

This result is consistent with our model, which predicts that the higher-energy ReCO terminus will be quenched by efficient PEnT and have a short-lived excited state, whilst the Rubpyamide acceptor terminus will possess a long-lived excited state.

**ps-TRIR Kinetics**
Figures 3.7.3.2 and 3.7.3.3 show the ps-TRIR kinetics (for bleach and transient where possible) obtained for the ReCO and Rubpyamide termini in different time domains in CH3CN solution. Figures 3.7.3.2 shows Δt = 0-50 ps and 3.7.3.3 shows Δt = 0-1 ns.

Δt = 0-50 ps
This kinetics recorded during this period show some evidence for PEnT taking place, as there is a fast decay process evident in both bleach and transient on the ReCO terminus (with an aggregate time constant of ca. 30 ps), accounting for ca. 20% of the total signal intensity. Unfortunately the data for the Rubpyamide terminus is not of
comparable clarity and there is no grow-in of absorbance which would be characteristic of PEnT, by comparison with previous examples in this Chapter (§3.5.3, 3.6.3). In this case the absorbances change very little with time – this could be due to the offsetting of grown-in processes in the Rubpyamide terminus by other photophysical processes (e.g. vibrational cooling or rapid excited state decay).

As discussed above, the heavy overlap of the amide bleach and transient bands must also be considered, as this may introduce artefacts into the spectra, and hence these artefacts may propagate into the kinetic traces presented. The donor data decay could correspond to a time constant for rapid PEnT in this supramolecular system. Analysis of the donor terminus data provides few conclusions as neither a clear grow-in or decay is observed over the $\Delta t = 0$-50 ps timescale.

$\Delta t = 0$-1 ns

During this time domain the processes described above are observed at early times following excitation, and these are followed by signal decay on the ReCO terminus and interestingly, a growth in absorbance of the bleach and transient signals for the Rubpyamide terminus. This indicates that PEnT may be occurring on this timescale more efficiently than over the $\Delta t = 0$-50 ps timescale, or that the processes which masked the observation of the grow-in at early times in the Rubpyamide terminus may no longer be occurring. These processes appear to be multi-exponential, with adequate fits employing 1 or 2 components giving good agreement with the data for the ReCO terminus, with time constants averaged over the bleach and transient traces of 50 ($\pm$35) ps and 1020 ($\pm$400) ps.
Figure 3.7.3.1: ps-TRIR spectra of ReCO-BL\textsuperscript{1}-Rubpyamide in CH\textsubscript{3}CN solution of ReCO (left) and Rubpyamide (right) termini for $\Delta t = 0$-2000 ps.
This may be due to the offsetting of excited state decay processes with PEnT-related grow-in of signal – or the opposite situation where there are multiple competing decay processes occurring at varying rates. There appears to be little consistency in the time constants determined for the fits to the experimental data, with time constants varying between 20 and 500 ps for the ReCO terminus. Good fits to the data could not be obtained for the Rubpyamide terminus, although it appears that the growth process occurs on a similar timescale to the long decay component on the ReCO terminus (i.e. 400-600 ps). This may represent a second time constant related to the rate of PEnT in this system, and may sample a different conformation or array of conformations than the ca. 30 ps time constant obtained earlier (vide supra).

ns-TRIR Spectroscopy and Kinetics
In CH₃CN our photophysical model predicts that the Rubpyamide terminus should have a sufficiently long-lived excited state to allow monitoring on the nanosecond timescale. Figures 3.7.3.4 and 3.7.3.5 show the spectroscopy and kinetics of ReCO-BL¹-Rubpyamide measured at Δt = 1-500 ns.

The ns-TRIR spectroscopy of the ReCO terminus following excitation at 355 nm is analogous to the ps-TRIR spectroscopy shown above, with three bleaches at 2023, 1916 and 1897 cm⁻¹ and three transient bands shifted to higher wavenumber at 2058, 1996 and 1959 cm⁻¹. The signals detected were weak, implying that most of the excited state had decayed on the picosecond timescale, although the signal intensities over the two experiments are unfortunately not directly comparable due to different laser excitation wavelengths, data collection times and sample concentrations. Similarly the ns-TRIR spectroscopy of the Rubpyamide terminus following excitation at 355 nm is analogous to the ps-TRIR results shown above, with an amide ν(CO) bleach centred at 1635 cm⁻¹ and a transient to lower wavenumber centred at 1625 cm⁻¹.
Figure 3.7.3.2: ps-TRIR kinetics of ReCO-BL$^1$-Rubpyamide in CH$_3$CN solution for the ReCO (transient at 2056 cm$^{-1}$, top left and bleach at 2024 cm$^{-1}$, bottom left) and Rubpyamide (transient at 1620 cm$^{-1}$, top right and bleach at 1636 cm$^{-1}$, bottom right) termini for $\Delta t = 0$-50 ps.
Figure 3.7.3.3: ps-TRIR kinetics of ReCO-BL$^1$-Rubpyamide in D$_2$O solution for the ReCO (transient at 2056 cm$^{-1}$, top left and bleach at 2024 cm$^{-1}$, bottom left) and Rubpyamide (transient at 1620 cm$^{-1}$, top right and bleach at 1636 cm$^{-1}$, bottom right) termini for $\Delta t = 0-2$ ns.
At first glance it is clearly evident from Figure 3.7.3.4 that there is a significant growth in the absorbance of both bleach and transient over the first few nanoseconds on the Rubpyamide terminus. Unfortunately the data spacing is insufficient to meaningfully subject it to a fitting routine, but it can be estimated that the time constant of this process in both cases is ca. 2-4 ns. This is well outside the time window for vibrational cooling and other primary photophysical processes but is in good agreement with the proposed timescale for PEnT from emission kinetics (Δt = 0-2 ns, §3.7.1) and this may be tentatively assigned to ReCO → Rubpyamide PEnT.

Analysis of the kinetics at the bleach and transient peak maxima of ReCO give biexponential fits, with the components having time constants of 1.9 (±0.6) and 12.4 (±3.8) ns. The 1.9 ns decay component matches well with the grow-in observed on the ReCO terminus and this strengthens the assignment of PEnT to this process.

There is some agreement between the values of the longer picosecond decay component and the shorter nanosecond decay component, suggesting that these are the same process. Analysis of the kinetics at the bleach and transient peak maxima of Rubpyamide also give biexponential fits, with components having average time constants of 7.5 (±3.7) and 210 (±76) ns. The latter could conceivably correspond to the natural excited state lifetime of the Rubpyamide excited state, with the difference in duration being explained by the higher concentrations and possibility of O₂ quenching associated with the TRIR experiment (see Chapter 6 for a more detailed discussion of this).

As with the investigations in §3.5 and 3.6 the processes above have been confirmed as intramolecular through experiments on solution mixtures of the model complexes which detected no PEnT processes.
Figure 3.7.3.4: ns-TRIR spectra of ReCO-BL²-Rubpyamide in CH₃CN solution for both termini for Δt = 0-300 ns.
Figure 3.7.3.5: ns-TRIR kinetics of ReCO-BL$^1$-Rubpyamide in CH$_3$CN solution for ReCO (transient at 2053 cm$^{-1}$, top left and bleach at 2024 cm$^{-1}$, bottom left) and Rubpyamide (transient at 1615 cm$^{-1}$, top right and bleach at 1634 cm$^{-1}$, bottom right) termini for $\Delta t = 0$-300 ns.
3.7.4. Computational Studies and Conformational Analysis

As discussed previously ReCO-BL\textsuperscript{1}-Rubpyamide has a considerable amount of conformational freedom due to the ability to freely rotate and tumble around the CH\textsubscript{2}CH\textsubscript{2} linkage of the BL\textsuperscript{1} ligand (§3.5.4). Molecular mechanics calculations have been performed using the SPARTAN02 Computational Suite to gain insight into the effect that this conformational freedom has on the parameters that govern the rate and efficiency of PEnT.

As with RuCN-BL\textsuperscript{1}-Rubpyamide, Four important cases of conformational behaviour have been isolated, as shown in Figure 3.7.4.1. These consist of two 'closed' structures where the two metal centres are in close proximity (although their ligands prevent closer approach) and two 'open' conformers in which the BL\textsuperscript{1} ligand has unfolded to form a near linear chain. The relative positions of the metal centres in these two pairs of conformers can be thought of as syn (i and iii) or anti (ii and iv) to each other. The intermetallic distances in these cases range from 5.5 Å in conformer (i) to 12.6 Å in conformer (iv). Therefore it is likely that an ensemble of conformers are present in solution with intermetallic distance 5.5 < r < 12.6 Å.

The electrostatic properties of ReCO-BL\textsuperscript{1}-Rubpyamide are different to the previous two dinuclear complexes discussed in this Chapter as there is no strong attraction between the two metal centres. The ReCO terminus is neutral and the Rubpyamide terminus is doubly cationic; thus the predisposition of ion-pairing postulated for RuCN-BL\textsuperscript{1}-Rubpyamide is unlikely to apply to ReCO-BL\textsuperscript{1}-Rubpyamide, and the relative proportions of these four conformers present in solution may well be different, leading to a different distribution of intermetallic distances and PEnT rates. This is obviously in addition to the significant difference in photophysics between the two systems, \textit{i.e.} the lack of a solvent-switchable change in the direction and gradient of PEnT.
Figure 3.7.4.1: Four important conformers of ReCO-BL$^1$-Rubpyamide calculated in the gas phase using SPARTAN's molecular mechanics package: i) and ii) are syn and anti folded conformers of closed structures; iii) and iv) are syn and anti conformers of open structures. Amide groups and $H$ atoms omitted for clarity.
3.7.5. ReCO-BL\textsuperscript{1}-Rubpyamide: Closing Points

The bimetallic complex ReCO-BL\textsuperscript{1}-Rubpyamide has been studied using UV/visible absorption and luminescence spectroscopy, cyclic voltammetry and TRIR spectroscopy in CH\textsubscript{3}CN solution.

UV/visible spectroscopy (§3.7.1) and cyclic voltammetry (§3.7.2) of ReCO-BL\textsuperscript{1}-Rubpyamide have shown that the ground state properties of the system are essentially a superposition of the constituent monometallic fragments most closely represented by the model complexes [ReCl(CO)\textsubscript{3}bpy] and Rubpyamide-BL\textsuperscript{1}. In an analogous manner to RuCN-BL\textsuperscript{1}-Rubpyamide, this supports our supposition that the saturated spacer connecting the two bpy ligands in the bridging ligand BL\textsuperscript{1} prevents electronic communication between the two termini.

Luminescence spectroscopy of ReCO-BL\textsuperscript{1}-Rubpyamide has shown that the spectral profile is dominated by emission arising from the Rubpyamide terminus, with emission being substantially more intense (up to an order of magnitude) than the emission measured (in either solvent) the RuCN-BL\textsuperscript{1}-Rubpyamide systems. This may be due to the large difference in the energies of the ReCO and Rubpyamide termini minimising any thermal deactivation pathways. TCSPC experiments have measured multiple time constants on the nanosecond timescale as in RuCN-BL\textsuperscript{1}-Rubpyamide, presumably due to the conformational freedom of the BL\textsuperscript{1} ligand giving rise to an ensemble of possible conformers which may be present in solution (§3.8.4, vide infra). Further luminescence lifetime measurements in the 0-10 ns time range successfully measured a grow-in of luminescence on the Rubpyamide terminus with a time constant of $\tau = 0.32$ (±0.04) ns. As the Rubpyamide terminus possesses the lowest energy $^3$MLCT excited state in the system and is the PEnT acceptor the observation of this grow-in apparently suggests that we are indeed observing the PEnT process directly.
TRIR spectroscopy has been used to directly monitor the processes occurring following photoexcitation of ReCO-BL\(^1\)-Rubpyamide on the picosecond and nanosecond timescales. As with RuCN-BL\(^1\)-Rubpyamide the ps and ns-TRIR data exhibits polyexponential time dependence, supporting the results from emission spectroscopy that a number of photophysical processes are occurring in solution. In agreement with the model presented in §3.7.3, the results presented in the same section provide some evidence for ReCO $\rightarrow$ Rubpyamide PEnT in CH\(_3\)CN in the $\Delta t = 0$-2 ns time range (Figure 3.7.3.3) as a clear grow-in of absorbance is observed in the spectra of the Rubpyamide (PEnT acceptor) terminus. The time constants for PEnT were not precisely determined due to the low instrumental spectral resolution of the PIRATE spectrometer (ca. 8-9 cm\(^{-1}\)) and noise fluctuations in the data but it may be visually estimated as $\tau_{\text{PEnT,CH}_3\text{CN}} \approx$ 0.5-3 ns. This value is in good agreement with the value obtained from luminescence spectroscopy ($\tau_{\text{PEnT}} \approx 0.3$ ns).

These PEnT time constants represent considerably lower PEnT rates than that determined for the RuCN-BL\(^2\)-Rubpyamide systems, presumably due to the difference in the energetics of the PEnT process. From \(^3\)MLCT energies, RuCN and Rubpyamide are quite closely matched in energy in high A. N. solvent environments with larger energy differences after the location of the lowest excited state reverses in lower A. N. environments ($\Delta E_{\text{D}_2\text{O}} \approx$ 200 cm\(^{-1}\) whereas $\Delta E_{\text{CH}_3\text{CN}} \approx$ 700 cm\(^{-1}\)). In contrast, ReCO and Rubpyamide have an energy difference of at least 750 cm\(^{-1}\) in CH\(_3\)CN.

A simple conformational analysis of ReCO-BL\(^1\)-Rubpyamide has been performed computationally (§3.7.4) and gave the structures of a number of important conformers of the flexible system, based on the metal centres being either syn or anti to each other and the structure adopting a folded or open geometry (Figure 3.7.4.1), as was the case with RuCN-BL\(^1\)-Rubpyamide. Due to the ease of rotation about the CH\(_2\)CH\(_2\) axis of BL\(^1\) it has been postulated that these structures are likely to be members of an ensemble of conformers present in solution, leading to a range of intermetallic distances between $5.5 < r < 12.6$ Å and presumably resulting in a wide variation in the
Forster PEnT rate. This provides a plausible explanation for the multiexponential kinetics obtained by luminescence and TRIR spectroscopy as interchromophoric distance and and predicted Forster PEnT rates are sensitively interdependent (§1.1.3).

The electrostatic properties of the system are different to those observed in the RuCN-BL\textsuperscript{x}-Rubpyamide systems; Rubpyamide is doubly cationic but ReCO is neutral and as a result ReCO-BL\textsuperscript{1}-Rubpyamide is not expected to undergo any ion-pairing in solution. This is likely to influence the conformer distribution of ReCO-BL\textsuperscript{1}-Rubpyamide present in solution when the spectroscopic experiments were conducted.

The luminescence and TRIR spectroscopic results obtained which have successfully observed PEnT have been presented and discussed in relation to the related RuCN-BL\textsuperscript{x}-Rubpyamide systems.

### 3.8. Discussion and Closing Points

In this Chapter a series of bimetallic complexes optimised for Förster (through-space) energy transfer have been presented and optical and vibrational spectroscopic measurements presented. This discussion will briefly compare the properties of the three bimetallic systems RuCN-BL\textsuperscript{1}-Rubpyamide, RuCN-BL\textsuperscript{2}-Rubpyamide and ReCO-BL\textsuperscript{1}-Rubpyamide highlighting the major differences and will discuss the findings of this study with the work of Simpson et al. on the related complex RuCN-BL-Rubpy.\textsuperscript{3}

One issue that has not been fully addressed thus far is that solvatochromism leads to changes in the absorption spectra of the two termini within the system, significantly for RuCN and less so for ReCO and Rubpyamide. For the TRIR experiments detailed in this Thesis laser excitation takes place at a single wavelength ($\lambda_{\text{ps-TRIR}} = 400$ nm and $\lambda_{\text{ns-TRIR}} = 355$ nm, see Chapter 6 for more details) and the solvent-based change in absorbance of the two termini at these wavelengths could lead to different proportions of the various excited states formed being populated in either solvent investigated. This effect adds a further degree of complexity to the unselective
excitation of either terminus by excitation at the two laser wavelengths employed in the TRIR spectra presented.

Secondly it should also be considered from an absorption cross-section perspective that the Rubpyamide terminus contains three "Ru $\rightarrow$ bpy" MLCT chromophores whilst RuCN and ReCO contain only one. Thus the 'relative oscillator strength' of the Rubpyamide terminus at 400 nm (near the absorption maximum for Ru $\rightarrow$ bpy $^1$MLCT transitions) may be considerably weighted towards a greater proportion of the excitation energy going into the Rubpyamide terminus in most cases. This may be particularly relevant in cases where Rubpyamide is the lower energy terminus (e.g. in D$_2$O for the RuCN-BL$^1$-Rubpyamide systems) as a majority of the excitation energy may be channelled into a chromophore which may not undergo PENt, thus affecting our ability to objectively probe the apparent efficiency of the PENt process.

A further point is that a Ru $\rightarrow$ bpyamide $^3$MLCT excited state may be delocalised over the two bpyamide ligands, whilst the Ru$_{CN} \rightarrow$ bpy and Re$_{CO}$ $\rightarrow$ bpy $^3$MLCT excited states must be localised. Such a 'localised-to-delocalised' transition may present additional energetic barriers or selection rules arising from changes in electronic structure and symmetry.$^{28}$

In many of the results presented in this Chapter the amount and quality of information which may be extracted from our TRIR spectra of in the v(CO)amide region have not been as diagnostic as was hoped when the experiments were designed. The reasons for this are not immediately clear but the large extent of overlap between bleach and transient bands may be a contributing factor. Additionally it is important to consider that the amide reporter group is reporting on the electron density of a ligand rather than directly reporting on the electronic structure of a metal centre.

It is difficult to objectively compare the quantitative values (e.g. excited state lifetimes) obtained in this Chapter due to the large number of variables and unknowns in the study, mainly arising from the complexity of the supramolecular systems.
presented, and this was a major reason why the Förster parameters of these systems has not been calculated. Good examples of this are the quantum yields for the bimetallic systems RuCN-BL\textsuperscript{1}-Rubpyamide and RuCN-BL\textsuperscript{2}-Rubpyamide. As the two systems have slightly different termini (the RuCN substitution patterns on the bpy ligand vary) their absorption spectra may not match closely and there may be differing degrees of solvatochromism and as a result the excited state energetics of the two systems (in a common solvent) may be different.

As a rule of thumb it appears that using the emission spectra with $\lambda_{ex} = 400$ nm may provide most insight as this wavelength is slightly higher in energy than the two ruthenium (II) based chromophores and absorbances are comparable in the two systems. The quantum yields for the two systems at excitation wavelengths of 400 ± 10 nm are within 10% of each other in D\textsubscript{2}O but when the RuCN terminus is at lower energy and luminescence quenching by PEnT is activated in CH\textsubscript{3}CN the values are somewhat different (within 50% of each other) but the differences are much more marked if it is compared with a luminescence spectrum measured with a lower energy $\lambda_{ex}$. This is presumably because the emission spectrum reports on two processes occurring in the system following excitation: bona fide PEnT and simple emission from the Rubpyamide terminus. Finding ways to disentangle these two processes, possibly though the use of luminescence polarisation anisotropy measurements, is an interesting avenue of investigation for future extensions to this work.

In the above case the large disparity in emission quantum yields in CH\textsubscript{3}CN of RuCN-BL\textsuperscript{2}-Rubpyamide (when compared with RuCN-BL\textsuperscript{1}-Rubpyamide) may be indicative of differences in the efficiency of the quenching process as related to the identity of the bridging ligand. A more detailed study of these systems with multiple bridging ligands would provide further insight into this hypothesis.

It is also important to consider the precise differences in $^3$MLCT excited state energy gradients between the two termini in each system (i.e. $\Delta E_{Donor} - \Delta E_{Acceptor}$) and how they vary with solvent. These energy gradients undoubtedly play a significant role in
controlling the dynamics of the PENt processes observed in this work (see Indelli et al., §3.1 for an illustration of this) and an extension of this work to study solvent mixtures or Acceptor Number media intermediate between D$_2$O and CH$_3$CN would be of great utility in probing this further.

Regarding the polyexponential nature of the luminescence and TRIR kinetic data presented throughout this Chapter, it is conceivable that in some cases there may be more than four time constants present in the TCSPC fits, which is a limitation of the spectrometer control software employed for their analysis. For example, the four exponent fit of TCSPC data of RuCN-BL$^{1}$-Rubpyamide has a $\chi^2$ value of $\sim$1.4. This suggests that the fit may not be statistically robust and that a greater number of time constants would be likely to improve the quality of the fit to the experimental data.

We must also attempt to contextualise this work in the frame of the proof of concept study carried out by Simpson et al. Our results are broadly comparable with theirs but this Thesis has revealed a level of depth and complexity to the primary photophysical processes occurring which was not observed in their initial study. There are a number of interesting differences between this work and their study which would provide subject matter for future work in this area.

Simpson et al. reported that the luminescence measured from RuCN-BL-Rubpy was monoexponential. This is very interesting since there is clearly a polyexponential time-dependence of all of the TCSPC traces and TRIR kinetics measured for bimetallic systems in this Chapter. A possible explanation for this may be related to conformational pre-organisation of the slightly different bridging ligand used in their study, based on a polyoxomethylene chain. Such chains are known to undergo conformational ordering in solution (with the chains aligned either anti or gauche with respect to each other depending on the solvent). This may introduce an energetic barrier to rotation which may reduce the conformational complexity of the system and thus result in a monoexponential TCSPC trace.
Additionally the longer chain length in RuCN-BL-Rubpy may also lead to a greater interchromophoric separation than in RuCN-BL'-Rubpyamide (14 Å c.f. 5-12 Å) which in turn may have lowered energetic barriers to rotation and increased conformational freedom. This may have the plausible consequence of "smearing" out the polyexponential time-dependence from 4/5/10 components to hundreds or even thousands of components and have it tend towards a mean value, which may be observed as a trace fitted with a monoexponential decay by the spectrometer control software.

The second set of differences concern the ligand set. The absence of amide groups in RuCN-BL-Rubpy means that the Ru_{bpy} → bpy and Ru_{bpy} → BL \(^3\)MLCT energies would be very similar and the possibility of delocalisation of an excited state over both these ligands is a possibility (§1.3). This ligand set would possibly prevent a 2-step PEnT process from occurring, leading to a simplification of the photophysics of the system and a smaller number of photophysical processes (and hence exponents) observed. Alternatively if the energy of (Ru_{bpy} → BL) < (Ru_{bpy} → bpy) then PEnT is a localised-localised process, contrasting with the delocalised-localised process in RuCN-BL'-Rubpyamide.

A comparison of Ward's systems with those in this Chapter support the notion that intermetallic distance may be a less suitable 'molecular yardstick' for PEnT and that the interchromophoric distance is preferable. This is due to the location of the MLCT charge density, spread over both the bipyridyl ligand and metal, which is the determining factor for the distance that such electron density must travel for PEnT to occur. If PEnT takes place between termini in Simpson's RuCN-BL-Rubpy (in which the two chromophores are 'proximal' to the bridging ligand \(i.e\.) on the near side of the bridging ligand to each other) then this represents a "direct line" whilst in this work due to the lower energy of the bpyamide ligands the excited state electron density is redistributed over more space, as the lowest energy MLCT molecular orbitals may either located proximal (RuCN) or distal (Rubpyamide) to the bridging ligand.
For the reasons above it would be inappropriate to give quantitative figures on the efficiency or rates of PFnT occurring in this Chapter as the detailed mechanisms and processes occurring appear too complicated to meaningfully apply the Förster treatment to. A further study involving a larger set of bichromophoric systems, a greater number of solvent environments and a great deal more experimental beamtime, preferably a combination of time-resolved UV/vis spectroscopy and TCSPC in conjunction with TRIR and Raman spectroscopy would be required to fully unravel the complexity of these systems.

The study in Chapter 4 discusses bimetallic complexes constructed from the same metal complex fragments but held together in very close proximity by a rigid, conjugated ligand. This will provide interesting material for comparison of similar metal complexes incorporated into two distinctly different classes of supramolecular system.
3.9. Appendix of Spectra and Voltammograms
Measurements on RuCNdmb in D$_2$O (black), CH$_3$CN (red) and CH$_3$OH (green). Left: UV/vis absorption spectra. The absorbances are normalised to the long-wavelength feature. Right: Uncorrected emission spectra recorded under identical conditions, with $\lambda_{\text{em}}$(H$_2$O) = 395 nm, $\lambda_{\text{em}}$(CH$_3$OH) = 420 nm and $\lambda_{\text{em}}$(CH$_3$CN) = 480 nm.

Measurements on RuCNdmb. Left: cyclic voltammogram in CH$_3$CN. Potentials are quoted vs. Fe/Fe$^\text{3+}$. The presence of features in the return wave of the reduction wave is due to decomposition of the solvent. Right: FTIR spectra of RuCNdmb in D$_2$O (black) and CH$_3$CN (red).
Rubpyamide-BL\textsuperscript{1}

Measurements on Rubpyamide-BL\textsuperscript{1} in D\textsubscript{2}O (black) and CH\textsubscript{3}CN (red). Left: UV/vis absorption spectra. The absorbances are normalised to the peak at 295 nm. Centre: Uncorrected emission spectra recorded under identical conditions $\lambda_{\text{em}}(\text{D}_2\text{O}) = 481$ nm and $\lambda_{\text{em}}(\text{CH}_3\text{CN}) = 468$ nm. Right: FTIR spectra in D\textsubscript{2}O (black) and CH\textsubscript{3}CN (red).

Cyclic voltammogram of Rubpyamide-BL\textsuperscript{1} in CH\textsubscript{3}CN. Potentials are quoted vs. Fc/Fc\textsuperscript{+}. 
Rubpyamide-BL$^2$

Measurements on Rubpyamide-BL$^2$ in D$_2$O (black) and CH$_3$CN (red). Left: UV/vis absorption spectra. The absorbances are normalised to the peak at 295 nm. Centre: Uncorrected emission spectra recorded under identical conditions $\lambda_{ex}(D_2O) = 476$ nm and $\lambda_{ex}(CH_3CN) = 474$ nm. Right: FTIR spectra in D$_2$O (black) and CH$_3$CN (red).
RuCN-BL$^1$-Rubpyamide

FTIR spectra of RuCN-BL$^1$-Rubpyamide in D$_2$O (left) and CH$_3$CN (right).

Cyclic Voltammogram of RuCN-BL$^1$-Rubpyamide in CH$_3$CN
RuCN-BL\textsuperscript{2}-Rubpyamide

Measurements on RuCN-BL\textsuperscript{2}-Rubpyamide in D\textsubscript{2}O (black) and CH\textsubscript{3}CN (red). Left: UV/vis absorption spectra. The absorbances are normalised to the peak at 295 nm. Centre: FTIR spectrum in D\textsubscript{2}O. Right: FTIR spectrum in CH\textsubscript{3}CN (red).

Cyclic Voltammogram of RuCN-BL\textsuperscript{2}-Rubpyamide in CH\textsubscript{3}CN
ReCO-BL$^1$-Rubpyamide

Measurements on ReCO-BL$^1$-Rubpyamide in CH$_3$CN. Left: UV/vis spectrum; right: FTIR spectrum.

Cyclic Voltammogram of ReCO-BL$^1$-Rubpyamide in CH$_3$CN
3.10. References


M. K. Kuimova, D. C. Grills, P. Matousek, A. W. Parker, X-Z. Sun, M. Towrie and M. W. George, 


Chapter 4:

Probing the Excited States of Ruthenium (II) and Rhenium (I) Complexes of 2,2’-Bipyrimididine using TRIR Spectroscopy
Chapter 4: Probing the Excited States of Ruthenium (II) and Rhenium (I) Complexes of 2,2'-Bipyrimidine using TRIR Spectroscopy

4.1. Introduction

4.1.1. 2,2'-Bipyrimidine As A Bridging Ligand

This chapter describes a study complementary to that discussed in Chapter 3, combining the same three d^6 transition metal building blocks fac-[ReCl(CO)_3], [Ru(CN)_4]^{2-} and [Ru(4,4'-CONEt_2)_{2}bpy]^{2+} with the potentially bis-chelating bridging ligand 2,2'-bipyrimidine (hereafter bpm). A series of 3 monometallic complexes, 3 homobimetallic systems (where the same d^6 unit is complexed to both sites of the bridging ligand) and 3 heterobimetallic systems (where different moieties are complexed to the bpm binding sites) have been designed and investigated using TRIR spectroscopy. This Chapter aims to provide insight into the efficacy, interplay and timescale of the photophysical processes occurring in such complexes in different solvent media; where possible the experiments have been conducted in D_2O and CH_3CN. This study provides a direct comparison of the ground and excited state electronic structure of complexes possessing metal centres held in close proximity with an estimated separation of ca. 5.5 Å,^1 by a conjugated bpm bridge (this Chapter) and a longer, saturated linkage terminating in two mono-chelating bipyridyl ligands resulting in a considerably larger intermetallic distance (up to ca. 13 Å, Chapter 3).

![Figure 4.1.1.1: The structures of 2,2'-bipyridine (bpy, left) and 2,2'-bipyrimidine (bpm, right)](image)

The salient properties of bpy and bpm will be compared before a discussion of previous studies of similar complexes in the scientific literature. Whilst bpy possesses two nitrogens which may form a chelating ring, bpm possesses four so that either
mono or bimetallic complexes may be prepared (Figure 4.1.1.1). At 298 K $d^6$ metal complexes of bpy may be treated as having a single low-lying $^3$MLCT excited state (see §1.4.1.1 for a discussion of this) sometimes with thermally accessible d-d transitions, giving rise to temperature-dependent luminescence properties. Spectroscopic studies have found that the electronic structure of bpm is more complicated than that of bpy, with a number of energetically close-lying orbitals making up the frontier orbital manifold (Figure 4.1.1.2). This may result in several electronic transitions being thermally accessible, i.e. at energies within $k_B T$ of the lowest $^3$MLCT excited state of bpm complexes. This is because the electronic selection rules which forbid transitions between these states in the free ligand are relaxed owing to the spin-orbit coupling effect of heavy metal atoms such as rhenium (I) and ruthenium (II) (§1.1).

![Figure 4.1.1.2: The electronic structure of frontier molecular orbitals of 2,2'-bipyrimidine. Adapted from Ref.\textsuperscript{2}](image)

A comparative study of the electrochemical properties of bpy and bpm reported that the two reductions of bpm (-2.42 and -2.94 V vs. Fc/Fc\textsuperscript{+}) occur more closely together and at less negative potentials than bpy (-2.56 and -3.20 V vs. Fc/Fc\textsuperscript{+}), presumably because of the presence of additional nitrogens in the bpm ring.\textsuperscript{3} This suggests that the
LUMOs in bpm are lower in energy and are in closer energetic proximity than the corresponding LUMO orbitals of bpy.

### 4.1.2. Relevant Studies of Re and Ru complexes of bpm

There have been several relevant studies incorporating the bpm ligand system into rhenium (I) and ruthenium (II) complexes, dating back to the first report of homoleptic and heteroleptic complexes of bpy and bpm by Sullivan et al. in 1977. This paper briefly reported the synthesis and properties of \([\text{Ru(bpm)}_3]^{2+}\), \([\text{Ru(bpy)}_2\text{(bpm)}]^{2+}\), \({\{(\text{Ru(bpy)}_2\text{(bpm})\}^4+}\) and \({\{(\text{Ru(bpy)}_2\text{(bpm})\}_3\text{Ru}^8+}\) using electronic absorption, emission and electrochemical properties in aqueous solution.

![Figure 4.1.2.1: The structures of \([\text{Ru(bpm)}_3]^{2+}\) (top left), \([\text{Ru(bpy)}_2\text{(bpm)}]^{2+}\) (top right), \({\{(\text{Ru(bpy)}_2\text{(bpm})\}^4+}\) (bottom left) and \({\{(\text{Ru(bpy)}_2\text{(bpm})\}_3\text{Ru}^8+}\) (bottom right).](image-url)
In comparison to [Ru(bpy)$_3$]$^{2+}$ additional low energy absorption bands in all the complexes were observed and assigned to Ru → bpm $^1$MLCT transitions, with these bands shifted to considerably lower energy in polynuclear complexes containing bridging bipyrimidine (Table 4.1.2.2). This observation was rationalised in terms of the bpm LUMO energy being lowered upon complexation of a second metal centre. Interestingly the homoleptic complex [Ru(bpm)$_3$]$^{2+}$ was found to emit strongly at room temperature (644 nm), with a quantum yield over double that of [Ru(bpy)$_3$]$^{2+}$ under the same conditions ($\phi = 0.09$ vs. 0.042) and its emission was found to be unquenched by the presence of O$_2$. The reason for this was not determined. The heteroleptic complex [Ru(bpy)$_2$(bpm)]$^{2+}$ was observed to emit weakly at 680 nm and the polynuclear systems {(Ru(bpy)$_2$)$_2$(μ–bpm)}$^{4+}$ and {(Ru(bpy)$_2$)(μ–bpm)$_2$Ru}$^{8+}$ were found to be non-luminescent at 298 K and the amount of luminescence detected at <10 K (between 725 and 850 nm) decreased as the nuclearity of the complexes increased, with no emission detected for the tetranuclear complex {(Ru(bpy)$_2$)(μ–bpm)$_3$Ru}$^{8+}$ under any conditions. Electrochemistry in H$_2$O was hampered by the narrow potential (ca. -1 to +1 V) so full characterisation was not possible. However it was noted that the replacement of bpy with bpm shifted the first (and second, if observed) reduction couples to considerably more positive potentials.

An article by Vogler and Kisslinger conducted a study of the rhenium (I) complexes fac-[ReCl(CO)$_3$(bpm)] and its dimer {(ReCl(CO)$_3$)$_2$(μ–bpm)} using absorption and emission spectroscopy in DMSO solution. They also reported the synthesis of the heterobimetallic d$^6$ complex {(ReCl(CO)$_3$)(μ–bpm)(Ru(bpy)$_2$)}$^{2+}$. The UV/vis absorption profiles of [ReCl(CO)$_3$(bpm)] were reported to be comparable to other complexes of α-diimine ligands ($\lambda_{max} = 371$ nm), whilst complexation of a second [ReCl(CO)$_3$] unit to form the homobimetallic complex gave rise to a broad, lower energy absorption band (with $\lambda_{max} = 450$ nm). The heterobimetallic complex possesses a very broad spectrum with discernible maxima at $\lambda_{max} = 408$ nm and 606 nm which were assigned to Ru → bpy and Ru → bpm $^1$MLCT transitions respectively (Table 4.1.2.2). The authors reasoned that the absorption spectrum was dominated by charge
transfer transitions arising from the ruthenium (II) fragment due to their greater oscillator strengths and that the Re → bpm features were present but unresolved.

Strong emission at 567 nm was reported for [ReCl(CO)₃(bpm)] in DMSO at room temperature, however the homobimetallic complex \{(ReCl(CO)₃)₂(μ-bpm)\} was found to be non-emissive under any conditions in contrast to other homobimetallic systems studied previously (vide supra). The heterobimetallic complex \{(ReCl(CO)₃)(μ-bpm)(Ru(bpYh))\}⁺⁺ was found to emit only at low temperature in the solid state at a wavelength (774 nm) similar to that reported by Sullivan et al. for \{(Ru(bpYh)₂(μ–bpm))\}⁺⁺. It was therefore postulated that the emitting excited state in this case was Ru → bpm ³MLCT. However the possibility of other charge-transfer transitions occurring in their bimetallic bpm-bridged systems was discussed in terms of an excited state deactivation pathway of the Re-Ru complex involving a low energy intervalence charge-transfer (IVCT) transition between the ReI_Ru III and ReII_Ru II redox states (Figure 4.1.2.2). Similarly, the lack of emission of the Re-Re complex was postulated to arise because of a low energy ReI-to-ReI IVCT transition, although no direct evidence was provided to substantiate this theory.

Sahai et al. published a study of Ru and Re complexes of bpm using FTIR, absorption and luminescence spectroscopy and electrochemistry. Their study re-examined many of the complexes presented by the groups of Meyer and Ludi and Vogler and Kisslinger (vide supra) but added valuable FTIR data (Table 4.1.2.1) for the rhenium (I) carbonyl complexes, and UV/Vis (Table 4.1.2.2) and spectroelectrochemical data (Table 4.1.2.3). The authors conducted an analysis of the relative energies of ¹MLCT and ³MLCT energies in these systems, suggesting that the energy gap between related singlet and triplet states decreases in the order: monometallic > homobimetallic > heterobimetallic, in concert with the decrease of the HOMO-LUMO energy gap. A luminescence lifetime of 76 ns was measured for [Ru(bpYh)(bpm)]⁺ in water. Additionally a discussion of the possible isomers of [(ReCl(CO)₃)₂(μ-bpm)] with Cl ligands oriented either syn or anti in relation to each other was explored – however it

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* This is inconsistent with other more recent studies, notably Kaim’s in CH₂Cl₂ (λ_em = 724 nm)⁶ and this work in CH₃CN at 298 K, 674 nm, vide infra.
was concluded that neither FTIR nor any of the other techniques employed in their study would be able to differentiate between the two possible stereoisomers.

\[
\text{(bpYhRul~bpm)ReIlCI(COhJ2}^{+} \\
\text{(bpYhRu"(bpm)ReICI(COhJ2}^{+}
\]

\[
\text{i)} \\
\text{ii)} \\
\text{iii):} \\
\text{iv)}
\]

**Figure 4.1.2.2: Photophysical scheme for \(\{(\text{ReCl(CO})_3(\mu-\text{bp})\text{Ru(bpY)}_2\}\}^{2+}\) in DMSO solution; i) and ii) are direct excitation, iii) is PEnT or IVCT and iv) is radiative/non-radiative relaxation to the ground state. Adapted from Ref.\textsuperscript{5}\)**

More recently Samuels and DeArmond have conducted a brief study of the spectroscopy, photophysics and electrochemistry of [Ru(CN)\(\text{a}(\text{bp})\text{m})\]\textsuperscript{2-}. As observed in other tetracyanoruthenate (II) complexes (§2.1), the MLCT absorption bands show considerable negative solvatochromism in CH\textsubscript{3}CH\textsubscript{2}OH and (CH\textsubscript{3})\textsubscript{2}NCHO (henceforth DMF) consistent with the effects of specific solvent-cyanide interactions (Table 4.1.2.2, §1.2.4). Cyclic voltammetry of [Ru(CN)\(\text{a}(\text{bp})\text{m})\]\textsuperscript{2-} showed one reversible reduction at \(E_{1/2} = -2.09\) V (vs. Fe/Fe\textsuperscript{+}) in DMF. This is less negative than the value obtained in the same study for [Ru(CN)\(\text{a}(\text{bp})\text{y})\]\textsuperscript{2-}, \(E_{1/2} = -2.46\) V.\textsuperscript{8}

<table>
<thead>
<tr>
<th>Complex</th>
<th>Solvent</th>
<th>FTIR Bands / cm\textsuperscript{-1}</th>
</tr>
</thead>
<tbody>
<tr>
<td>[ReCl(CO)(\text{a}(\text{bp})\text{y})]]</td>
<td>CH\textsubscript{2}Cl\textsubscript{2}</td>
<td>2024 1921 1899</td>
</tr>
<tr>
<td>[ReCl(CO)(\text{a}(\text{bp})\text{m})]]</td>
<td>Solid</td>
<td>2033 1906 1899</td>
</tr>
<tr>
<td>({(\text{ReCl(CO})_3(\mu-\text{bp})})</td>
<td>Solid</td>
<td>2028 1908</td>
</tr>
<tr>
<td>({(\text{ReCl(CO})_3(\mu-\text{bp})\text{Ru(bpY)}_2}}^{2+})</td>
<td>Solid</td>
<td>2034 1920 1915</td>
</tr>
</tbody>
</table>

**Table 4.1.2.1: FTIR data for rhenium (I) bipyrimidine complexes. Adapted from Refs.\textsuperscript{7,9}\**
4.1.2.2: Relevant UV/vis data for ruthenium (II) and rhenium (I) bipyrimidine complexes. Adapted from Refs. 4,5,7,9

<table>
<thead>
<tr>
<th>Complex</th>
<th>Solvent</th>
<th>λ_{obs} / nm</th>
</tr>
</thead>
<tbody>
<tr>
<td>[Ru(bpy)_3]^{2+}</td>
<td>H₂O</td>
<td>452</td>
</tr>
<tr>
<td>[Ru(bpm)_3]^{2+}</td>
<td>H₂O</td>
<td>452</td>
</tr>
<tr>
<td>[Ru(bpy)_2(bpm)]^{2+}</td>
<td>CH₃CN</td>
<td>480 (sh)</td>
</tr>
<tr>
<td>{(Ru(bpy)_2)(μ-bpm)}^{4+}</td>
<td>CH₃CN</td>
<td>594</td>
</tr>
<tr>
<td>[ReCl(CO)_3(bpm)]</td>
<td>CH₃CN</td>
<td>384</td>
</tr>
<tr>
<td>{(ReCl(CO)_3)(μ-bpm)(Ru(bpy)_2)}^{2+}</td>
<td>CH₃CN</td>
<td>606 (sh)</td>
</tr>
<tr>
<td>[Ru(CN)_4(bpm)]^{2-}</td>
<td>(CH₃)₂NCHO</td>
<td>590</td>
</tr>
<tr>
<td>[Ru(CN)_4(bpm)]^{2-}</td>
<td>CH₃OH</td>
<td>490</td>
</tr>
</tbody>
</table>

Table 4.1.2.2: Relevant UV/vis data for ruthenium (II) and rhenium (I) bipyrimidine complexes. Adapted from Refs. 4,5,7,9

4.1.3. Relevant Studies of Other Polynuclear Complexes

There have been a number of time-resolved spectroscopic studies of bimetallic Ru and Re complexes in solution which have investigated electronic structure and dynamics in excited states following excitation into the MLCT manifold of such complexes. The aim of these studies is usually to gain insight into excited state electronic structure and
the extent and timescale of excited state localisation/delocalisation. A particularly relevant example of this is the study of a series of bimetallic complexes constructed from [ReCl(CO)₃] units bridged by different polypyridyl ligands (Figure 4.1.3.1a).¹⁰ These complexes were studied using ps-TRIR spectroscopy in the ν(CO) spectral region and as such provide good reference data for the carbonyl spectra presented in the remainder of this Chapter. Similar TRIR spectra were obtained for all complexes, with the three ν(CO) bands observed in the ground state FTIR spectra being split into two sets of three bands, with one set shifted to higher wavenumber than the bleaches and the other set shifted to lower wavenumber (Figure 4.1.3.1b). This was found to be consistent with ‘asymmetry’ in the electronic structure of the excited state, with the species being formed best described as a localised excited state on the IR timescale (§1.3) with the form [Re(BL⁻)Re⁺].

![Figure 4.1.3.1: a) The bridging ligands used in the TRIR study (top) and an example of the complexes prepared (bottom); b) TRIR spectra of [Re(CO)₃Cl]₂BL1 in CH₂Cl₂ solution following 606 nm excitation (top) and fits to the experimental bands separated into three sets of three. Adapted from Ref.¹⁰](image-url)
There have been a number of more recent studies of bpm complexes of ruthenium (II), rhenium (I) and other photophysically interesting transition metals such as platinum (II),\textsuperscript{11} but none of them have adequately answered the key questions posed by the studies discussed above:

- Why do rhenium (I) and ruthenium (II) complexes of 2,2'-bipyrimidine exhibit much weaker luminescence and shorter lifetimes than the corresponding 2,2'-bipyridine complexes?

- What is the electronic structure of the excited states formed upon photoexcitation of d\textsuperscript{6} metal complexes of bpm, and how do they evolve after formation? Are they localised or delocalised on the IR timescale? Can vibrational spectroscopy be used to directly monitor these states more effectively than optical spectroscopy?

- Some bpm-bridged bimetallic systems may have several different MLCT chromophores. Considering this in combination with unselective excitation due to broad absorption bands, the rich electronic manifold of the bpm ligand and other IL/MC transitions and multiple vibrational spectroscopic signals from reporter ligands will the ultrafast dynamics of these systems be amenable to study by vibrational spectroscopy?

- Can IVCT processes occur in mixed valence homobimetallic or heterobimetallic bpm complexes? Can they be observed using vibrational spectroscopy?

In the remainder of this Chapter the results from spectroscopic and electrochemical studies will aim to address these questions.
4.2. \textit{fac-[ReCl(CO)\textsubscript{5}(bpm)]}, ReBPM

\begin{center}
\includegraphics[width=0.2\textwidth]{ReBPM.png}
\end{center}

\textbf{ReBPM} has been studied before using absorption, emission and FTIR spectroscopy and electrochemistry, as discussed above (§4.1.2). We have conducted these experiments in addition to IR OTTLE and TRIR experiments in order to study the electronic structure of its ground state, electrochemically generated species and its lowest-lying electronic excited state.

\textbf{ReBPM} is insoluble in aqueous media and has only been studied in CH\textsubscript{3}CN. The complex is not significantly solvatochromic and thus its photophysical and electrochemical properties are expected to vary little with solvent medium.

4.2.1. UV/visible Absorption and Luminescence Spectroscopy

The UV/visible absorption spectrum of \textbf{ReBPM} has been measured in CH\textsubscript{3}CN (see Appendix 4.12 for spectra). The results are presented in Table 4.2.1.1 and closely match literature values (Table 4.1.2.2).\textsuperscript{5}

<table>
<thead>
<tr>
<th>Solvent</th>
<th>$\lambda_{\text{max}}$ / nm</th>
</tr>
</thead>
<tbody>
<tr>
<td>CH\textsubscript{3}CN</td>
<td>&lt; 300</td>
</tr>
</tbody>
</table>

<table>
<thead>
<tr>
<th>Assignment</th>
<th></th>
<th></th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td>IL $\pi-\pi^*$ (bpm)</td>
<td>$^1$MLCT (Re)</td>
<td>$^1$MLCT (Re)</td>
<td></td>
</tr>
</tbody>
</table>

\textit{Table 4.2.1.1: Table of peak positions and assignments for the UV/visible absorption spectrum of ReBPM in CH\textsubscript{3}CN solution.}
The luminescence spectrum of \textbf{ReBPM} was measured in CH$_3$CN solution (see Appendix 4.12 for spectra). Excitation into the Re \(\rightarrow\) bpm MLCT manifold at 384 nm results in weak luminescence with a broad, structureless profile centred at 674 nm, in contrast to the results of Vogler’s study which reports an emission maximum of 567 nm.\(^5\) It is conceivable that a feature due to Raman scattering from solvent vibrations may be responsible for Vogler’s different interpretation of the emission spectrum. The luminescence is significantly red-shifted in comparison to the spectrum of [ReCl(bpy)(CO)$_3$], which is centred at 612 nm.\(^12\) This has been rationalised in terms of a smaller HOMO-LUMO energy gap in bpm-containing complexes compared to bpy analogues as discussed in §4.1. In both cases the emission intensity was insufficient to perform TCSPC experiments to determine the lifetime.

\subsection*{4.2.2. Cyclic Voltammetry}

Cyclic voltammetry of \textbf{ReBPM} in CH$_3$CN was performed (see Appendix 4.12 for voltammograms) and the results are presented in Table 4.2.2.1 below. Features corresponding to one reversible reduction couple and one irreversible oxidation couple are observed. The results are comparable to previous studies (Table 4.1.2.3) which also detect a second (irreversible) redox couple at more negative potentials, corresponding to a second reduction located on the bpm ligand.\(^7\)

\begin{table}[h]
\centering
\begin{tabular}{ccc}
\hline
$E_{\text{oc}}$/V (vs. Fc/Fc$^+$) & Reversibility & Assignment \\
\hline
+1.05 & Irreversible & Re$^{\text{III}}$ \\
-1.43 & Reversible & (bpm)$^{0/\text{e}^-}$ \\
\hline
\end{tabular}
\caption{Table of potentials for redox couples and assignments from the cyclic voltammogram of ReBPM in CH$_3$CN solution.}
\end{table}
ReBPM has been studied using FTIR, IR OTTLE and ps-TRIR spectroscopy in CH₃CN solution. Figure 4.2.3.1 shows the FTIR and ps-TRIR spectra obtained and Figure 4.2.3.2 shows the spectra obtained from the IR-OTTLE experiment for ReBPM in CH₃CN solution. Figure 4.2.3.3 shows the ps-TRIR kinetics obtained for bleach and transient bands in CH₃CN.

Acetonitrile Solution

**FTIR Spectroscopy**

The FTIR spectrum of ReBPM in CH₃CN solution in the v(CO) region has been measured and is shown in Figure 4.2.3.1 (bottom). In the carbonyl region there is a peak profile characteristic of the rhenium (I) tricarbonyl unit with 3 resolved maxima at 2028, 1924 and 1909 cm⁻¹. This is in good agreement with previous studies in the literature (Table 4.1.2.1).

**IR OTTLE Spectroelectrochemistry**

FTIR spectra were taken as the complex underwent its first reduction; the difference spectrum obtained by subtracting the starting FTIR from the final spectrum of the 1-electron reduced species is shown in Figure 4.2.3.2. Newly formed bands are evident at 2007 and 1878 cm⁻¹, which are shifted to lower wavenumber to the bands of the neutral species at 2028, 1924 and 1909 cm⁻¹. This is consistent with an increase in electron density at the metal due to the reduction of the bipyrimidine ligand.

**ps-TRIR Spectroscopy**

The ps-TRIR spectral profile obtained following 400 nm excitation contains three bleaches at 2028, 1924 and 1909 cm⁻¹. Three transient bands are detected at higher energy to the bleaches, which narrow and blue-shift over the first 20 ps due to vibrational cooling and settle at 2079, 2002 and 1965 cm⁻¹. This is consistent with a
decrease in electron density at the metal centre as a result of the formation of a \( \text{Re} \rightarrow \text{bpm}^{3}\) MLCT excited state.

**Figure 4.2.3.1:** The FTIR (bottom) and ps-TRIR spectra between 1 and 2500 ps (top) of ReBPM in CH\textsubscript{3}CN solution.

**Figure 4.2.3.2:** The FTIR (bottom) and IR OITLE difference spectra [ReBPM\textsuperscript{-} minus ReBPM] (top) of ReBPM in CH\textsubscript{3}CN solution.
The magnitude of the band shift between ground and excited state spectra of ReBPM ($\Delta_{ES-GS} = 50-80 \text{ cm}^{-1}$) is slightly larger than that of [ReCl(bpy)(CO)$_3$] ($\Delta_{ES-GS} = 40-66 \text{ cm}^{-1}$), possibly indicating that more charge is transferred to the bpm ligand as well as indicating that the HOMO-LUMO energy gap is different in the two complexes.\textsuperscript{9,13} There is considerable overlap between the bleaches and their associated transient bands. A monoexponential decay is observed following vibrational cooling, with $\tau = 1170 (\pm 72)$ ps (averaged over both bleach and transient features). This is significantly shorter then the excited state lifetime of [ReCl(bpy)(CO)$_3$]\textsuperscript{12,14} and consistent with the 'energy-gap law'\textsuperscript{15} and the lack of any significant emission from this complex.

4.2.4. ReBPM: Closing Points

The lowest-lying excited state of ReBPM is assigned to Re $\rightarrow$ bpm $^3$MLCT using TRIR spectroscopy, as previously reported in the literature. Upon comparison of the
results with \([\text{ReCl(bpy)(CO)}_3]\) it is evident that \textbf{ReBPM} possesses lower energy absorption and emission bands, a more accessible 1\textsuperscript{st} reduction, a similar FTIR peak profile and TRIR spectrum, a shorter 3MLCT excited state lifetime and lower quantum yield. These results support the notion that bpm-containing complexes exhibit smaller HOMO-LUMO energy gaps than their bipyridine analogues. Interesting comparisons with the homobimetallic complex (\textbf{Re}_2\text{BPM} \S 4.5) will give insight into the effect of complexing a second metal to the bpm ligand.

\begin{equation}
X_2[\text{Ru(CN)}_4(bpm)](X = \text{K, PPN}), \text{RuCNBPM}
\end{equation}

The spectroscopy and electrochemistry of \textbf{RuCNBPM} have been presented previously in a comparison of tetracyanoruthenate (II) complexes \(\S 2.2\) and the results will be discussed here in more detail. As with all other complexes containing the \([\text{Ru(CN)}_4]^2-\) unit, strong negative solvatochromism is observed due to SSDA interactions between the cyanide ligands and solvent molecules. As a result \textbf{RuCNBPM} exhibits significantly different properties in D\textsubscript{2}O and CH\textsubscript{3}CN.

\textbf{4.3.1. UV/visible Absorption and Luminescence Spectroscopy}

The UV/visible absorption spectra of \textbf{RuCNBPM} have been measured in CH\textsubscript{3}CN and D\textsubscript{2}O (see Appendix 4.12 for spectra) and have been discussed earlier (see \S 2.2.1). Strong negative solvatochromism is clearly evident in the differences in peak centre between the two solvents and is typical for complexes of this type. The results are
summarised in Table 4.3.1.1 and agree well with a previous study (Table 4.1.2.2). No luminescence from RuCNBPM was detected at 298 K in either D$_2$O or CH$_3$CN.

<table>
<thead>
<tr>
<th>Solvent</th>
<th>$\lambda_{max}$ / nm</th>
</tr>
</thead>
<tbody>
<tr>
<td>D$_2$O</td>
<td>&lt;300 345 440</td>
</tr>
<tr>
<td>CH$_3$CN</td>
<td>&lt;300 416 592</td>
</tr>
<tr>
<td>Assignment</td>
<td>IL $\pi$-$\pi^*$ (bpm)</td>
</tr>
</tbody>
</table>

Table 4.3.1.1: Table of peak positions and assignments for UV/visible absorption spectra of RuCNBPM in D$_2$O and CH$_3$CN solution.

4.3.2. Cyclic Voltammetry

Cyclic voltammetry of RuCNBPM was measured in D$_2$O (see Appendix 4.12 for voltammograms) and the results are presented in Table 4.3.2.1 below. Features corresponding to one irreversible oxidation couple was observed – the reduction couples are ostensibly outside the narrow D$_2$O potential window (-1 to +1 V). The results are unfortunately incomparable with literature data (§4.1) due to differences in solvent; Samuels and DeArmond used DMF, a low A. N. solvent whereas D$_2$O was used here.

<table>
<thead>
<tr>
<th>$E_{1/2}$ /V (vs. Fc/Fc$^+$)</th>
<th>Reversibility</th>
<th>Assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>$E_p = +1.45$</td>
<td>Irreversible</td>
<td>Ru$^{III}$</td>
</tr>
</tbody>
</table>

Table 4.3.2.1: Table of potentials for redox couples and assignments from the cyclic voltammogram of RuCNBPM in D$_2$O solution.

4.3.3. FTIR, IR OTTLE and Time-Resolved Infrared Spectroscopy

RuCNBPM has been studied using ps and ns-TRIR spectroscopy in D$_2$O and CH$_3$CN solution. IR OTTLE spectroelectrochemistry in CH$_3$CN was unsuccessful as the 1st reduction was found to be chemically irreversible on the timescale of the experiment.
Deuterium Oxide Solution

Figure 4.3.3.1 shows the FTIR and ns-TRIR spectra of RuCNBPM measured in D₂O solution. Figure 4.3.3.2 shows the ns-TRIR kinetics obtained for bleach and transient bands in D₂O.

The FTIR spectrum of RuCNBPM in D₂O solution in the ν(CN) region has a peak profile characteristic of the tetracyanoruthenate (II) unit with 3 resolved maxima at 2095, 2057 and 2043 cm⁻¹. The ns-TRIR spectral profile of RuCNBPM following 355 nm excitation features a broad bleach with components corresponding to the band positions in the FTIR and a transient band at higher wavenumber at 2105 cm⁻¹. This is consistent with a decrease in electron density at the metal centre as a result of Ru → bpm ³MLCT excited state formation. A monoexponential decay is observed with τ = 3.4 (±0.3) ns (averaged over both bleach and transient features). Although the spectral profile is similar, the lifetime is significantly shorter than that of [Ru(CN)₄(bpy)]²⁻ (§2.1) and consistent with the lack of detected emission from this complex.

Figure 4.3.3.1: The FTIR (bottom) and ns-TRIR spectra between 1 and 50 ns (top) of RuCNBPM in D₂O solution.
Figure 4.3.3.2: Single-pixel kinetics of bleach (2058 cm\(^{-1}\), bottom) and transient (2106 cm\(^{-1}\), top) ns-TRIR bands of RuCNBPM in D\(_2\)O solution.

Acetonitrile Solution

Figure 4.3.3.3 shows the FTIR and ps-TRIR spectra of RuCNBPM measured in CH\(_3\)CN solution. Figure 4.3.3.4 shows the ps-TRIR kinetics obtained for bleach and transient bands in CH\(_3\)CN.

The FTIR spectrum of RuCNBPM in CH\(_3\)CN solution in the v(CN) region has a peak profile characteristic of the tetracyanoruthenate (II) unit with 4 resolved bands at 2092, 2074, 2068 and 2059 cm\(^{-1}\). The ps-TRIR spectral profile obtained following 400 nm excitation consists of a broad bleach with components corresponding to the band positions in the FTIR and a transient band at higher wavenumber centred at 2098 cm\(^{-1}\). As in D\(_2\)O this is consistent with a decrease in electron density at the metal centre as a result of Ru → bpm \(^3\)MLCT excited state formation.
Figure 4.3.3.3: The FTIR (bottom) and ps-TRIR spectra between 1 and 1000 ps (top) of RuCNBPM in CH$_3$CN solution.

Figure 4.3.3.4: Single-pixel kinetics of bleach (2064 cm$^{-1}$, bottom) and transient (2099 cm$^{-1}$, top) ps-TRIR bands of RuCNBPM in CH$_3$CN solution.
A monoexponential decay is observed with $\tau = 249 \ (\pm 51) \ \text{ps}$ (averaged over both bleach and transient features). This is an order of magnitude shorter than the excited state lifetime of RuCNBPM in D$_2$O, which can be attributed to two principal reasons. Firstly, the attenuation of decay pathways in deuterated solvents$^{17}$ (see §2.1) and secondly the strong solvatochromism affecting the HOMO-LUMO energy gap and hence the lifetime in accordance with the 'energy-gap law'.$^{15}$

4.3.4. RuCNBPM: Closing Points

RuCNBPM exhibits strong negative solvatochromism in a similar manner to other tetracyanoruthenate complexes studied. In comparison to its bipyridine analogues [Ru(CN)$_4$(bpy)]$^{2-}$ and RuCNdmb it exhibits lower energy absorption bands, a dramatic decrease in emission (such that it is undetected) and a significantly shorter excited state lifetime. TRIR spectroscopy in the v(CN) spectral region facilitates direct monitoring of the excited states formed on the picosecond and nanosecond timescales. These observations agree with previous results showing that the HOMO-LUMO energy gap is significantly smaller in [Ru(CN)$_4$(NN)]$^{2-}$complexes containing bipyridimide than other polypyridyl ligands investigated in this Thesis (see Chapter 2).

4.4. [Ru((4,4'-CONEt$_2$)$_2$bpy)$_2$(bpm)]Cl$_2$, RubpyamideBPM
RubpyamideBPM is similar to the RubpyamBL\textsuperscript{x} complexes studied in Chapter 3, in that it possesses two \((4,4'-\text{CONEt}_2)\text{bpy})\) (henceforth known as bpyamide) ligands and a third different bidentate polypyridyl ligand with vacant coordination sites available for complexation of a second metal. As there are two types of chromophores in the complex - Ru \(\rightarrow\) bpm and Ru \(\rightarrow\) bpyamide MLCT - the photophysical behaviour of the complex may be complicated; in addition to two types of possible \(^3\text{MLCT}\) excited states, there may be the possibility of interconversion between them on the ultrafast timescale. As with the RubpyamBL\textsuperscript{x} complexes there is little solvatochromism expected and as a result the properties of RubpyamideBPM are predicted to be similar in both \(\text{D}_2\text{O}\) and \(\text{CH}_3\text{CN}\).

4.4.1. UV/visible Absorption and Luminescence Spectroscopy

The UV/visible absorption spectra of RubpyamideBPM have been measured in \(\text{D}_2\text{O}\) and \(\text{CH}_3\text{CN}\) (see Appendix 4.12 for spectra) and the results are summarised in Table 4.4.1.1. The absorption spectra of these complexes are only slightly affected by solvent, with differences of a few nanometres wavelength between related absorption peaks in \(\text{D}_2\text{O}\) and \(\text{CH}_3\text{CN}\). More low energy absorption bands are observed in comparison with the other monometallic bpm complexes studied here (§4.2.1, §4.3.1), due to the presence of Ru \(\rightarrow\) bpyam \(^1\text{MLCT}\) transitions in addition to the Ru \(\rightarrow\) bpm \(^1\text{MLCT}\) transitions. There are two transitions at ca. 420 and 360 nm, whose positions do not vary significantly with either solvent or nuclearity (when analysed in conjunction with \((\text{Rubpyamide})_2\text{BPM}\), §4.7.1) and hence are tentatively assigned to the Ru \(\rightarrow\) bpyam \(^1\text{MLCT}\) transitions. The bands at ca. 450 and 330 nm in 3 are tentatively assigned to two Ru \(\rightarrow\) bpm \(^1\text{MLCT}\) transitions, consistent with the presence of two different low-lying \(\pi^*\) bpm orbitals giving two \(^1\text{MLCT}\) transitions to bpm. The spectral profiles obtained are reasonably similar to those reported for \([\text{Ru(bpy)}_2(\text{bpm})]^2+\) (Table 4.1.2.2).\footnote{7}

The luminescence of RubpyamideBPM in \(\text{CH}_3\text{CN}\) (see Appendix 4.12 for spectra) is relatively intense for a bpm-containing complex, with a broad structureless emission...
profile centred at 648 nm which decays with a time constant of 116 (±10) ns. The quantum yield is however an order of magnitude smaller than the corresponding Rubpyamide-BL (Chapter 3) complexes, with \( \phi = 0.001 \). This comparatively weak luminescence is in agreement with published results for \([\text{Ru}(\text{bpy})_2(\text{bpm})]\)^{2+}.7

<table>
<thead>
<tr>
<th>Solvent</th>
<th>( \lambda_{\text{max}} / \text{nm} )</th>
</tr>
</thead>
<tbody>
<tr>
<td>D_2O</td>
<td>&lt; 300 330 358 424 448</td>
</tr>
<tr>
<td>CH_3CN</td>
<td>&lt; 300 330 358 418 453</td>
</tr>
<tr>
<td>Assignment</td>
<td>IL ( \pi-\pi^* ) &lt;br&gt; (bpm)</td>
</tr>
</tbody>
</table>

Table 4.4.1.1: Table of peak positions and assignments for UV/visible absorption spectra of RubpyamideBPM in D_2O and CH_3CN solution.

4.4.2. Cyclic Voltammetry

Cyclic voltammetry of RubpyamideBPM was measured in CH_3CN (see Appendix 4.12 for voltammograms) and the results are presented in Table 4.4.2.1 below. Features corresponding to three reversible reductions and one irreversible oxidation couple are observed. The redox potentials and assignments are comparable to studies of RubpyamideBL in §3.3.2, with the notable difference of the more facile reduction of bpm than bipyridine-based ligands.

<table>
<thead>
<tr>
<th>( E_{\text{Vs}} / V ) (vs. Fc/Fc(^+))</th>
<th>Reversibility</th>
<th>Assignment</th>
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<tr>
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<td>-1.30</td>
<td>Reversible</td>
<td>(bpm)^{0/-}</td>
</tr>
<tr>
<td>-1.60</td>
<td>Reversible</td>
<td>Amide reduction</td>
</tr>
<tr>
<td>-1.81</td>
<td>Reversible</td>
<td>Amide reduction</td>
</tr>
</tbody>
</table>

Table 4.4.2.1: Table of potentials for redox couples and assignments from the cyclic voltammogram of RubpyamideBPM in CH_3CN solution.
4.4.3. FTIR, IR OTTLE and Time-Resolved Infrared Spectroscopy

RubpyamideBPM has been studied using ps and ns-TRIR spectroscopy in D$_2$O and CH$_3$CN solution. An IR OTTLE experiment to monitor the changes in band profile as the complex underwent its 1$^{st}$ reduction in CH$_3$CN (not shown) did not detect any significant spectral changes, ostensibly due to the lack of electronic coupling between the bipyrimidine ligand and the amide substituents on the bpy ligands.

**Deuteration Oxide Solution**

Figures 4.4.3.1 and 4.4.3.3 show the FTIR, ps-TRIR and ns-TRIR spectra obtained of RubpyamideBPM in D$_2$O solution. Figures 4.4.3.2 and 4.4.3.4 show the ps and ns-TRIR kinetics obtained for bleach and transient bands in D$_2$O.

The FTIR spectrum of RubpyamideBPM in D$_2$O solution in the amide v(CO) region has a single amide peak present at 1617 cm$^{-1}$ and a weak band at 1580 cm$^{-1}$ which is assigned to a vibrational mode of bipyrimidine.

![Figure 4.4.3.1: The FTIR (bottom) and ps-TRIR spectra between 1 and 1000 ps (top) of RubpyamideBPM in D$_2$O solution.](image-url)
Following 400 nm excitation in D$_2$O, the ps-TRIR spectrum of RubpyamideBPM exhibits several features. The parent v(CO) amide band at 1617 cm$^{-1}$ is bleached and transient bands at lower and higher energy is observed. This has been confirmed through scaled subtraction and fitting analysis (not shown). The bleach band appears to grow in with a time constant of 7 (±2) ps; this is thought to be an artefact due to heavy overlap with associated transient bands (vide infra) which imprint their dynamics on the single-pixel bleach trace. The higher energy transient band is centred at 1636 cm$^{-1}$ and grows in with a time constant of 7 (±2) ps, whilst the lower energy band is centred at 1593 cm$^{-1}$ and decays with a time constant of 6 (±2) ps. The two transient bands correspond to the two possible excited states. Initial excitation is non-selective (cf. the electronic absorption spectrum, vide supra) so both [(bpyam)(bpyam$^*$)Ru$^{III}$(bpm)]$^{2+}$ (hereafter $^3$MLCT(bpyam)) and [(bpyam)$_2$Ru$^{III}$(bpm$^*$)]$^{2+}$ (hereafter $^3$MLCT(bpm)) are formed following excitation.
The transient band at higher energy corresponds to a decrease in electron density on the bpyamide ligands which is consistent with the formation of a $^3$MLCT(bpm) excited state. Conversely the lower energy transient band corresponds to an increase in electron density on the bpyamide ligands that accompany the formation of a $^3$MLCT(bpyam) excited state. These bands can therefore be used to monitor the formation and decay of the $^3$MLCT(bpm) excited state. The grow-in of the 1593 cm$^{-1}$ band corresponds to formation of the $^3$MLCT(bpm) state from the higher-energy $^3$MLCT(bpyam) excited state, i.e. relaxation of the higher-lying bpyam-based to the lower-lying bpm-based $^3$MLCT state occurs with a half-life of ca. 6 ps. Following this excited state interconversion there is no significant change in spectral profile on the picosecond timescale.

ns-TRIR spectra of RubpyamideBPM in D$_2$O were recorded using excitation at 355 nm. The synchronous decay of the 1593 and 1617 cm$^{-1}$ bands was observed with a mean time constant of 59 (±1) ns. The bipyrimidine bleach at 1580 cm$^{-1}$ and its associated transient at 1560 cm$^{-1}$ decay with the same (60 ns) lifetime as the higher energy amide transient band as they are also monitoring the decay of the $^3$MLCT(bpm) excited state. A rationalisation of the difference in lifetimes measured by TRIR and luminescence spectroscopy is presented in the Experimental section of this Thesis (Chapter 6).

**Acetonitrile Solution**

Figures 4.4.3.5 and 4.4.3.6 show the FTIR, ps-TRIR and ns-TRIR spectra obtained of RubpyamideBPM in CH$_3$CN solution. Figures 4.4.3.7 and 4.4.3.8 shows the ps and ns-TRIR kinetics obtained for bleach and transient bands in CH$_3$CN. The FTIR spectrum of RubpyamideBPM in CH$_3$CN solution in the amide $\nu$(CO) region has a single amide peak present at 1641 cm$^{-1}$ and a weak band at 1581 cm$^{-1}$ which is assigned to a vibrational mode of bipyrimidine. As in D$_2$O, ps-TRIR spectroscopy of RubpyamideBPM in CH$_3$CN gives a bleach of the main $\nu$(CO) amide band at 1641 cm$^{-1}$ following 400 nm excitation.
Figure 4.4.3.3: The FTIR (bottom) and ns-TRIR spectra between 1 and 750 ns (top) of RubpyamideBPM in D2O solution.

Figure 4.4.3.4: Single-pixel kinetics of bleach (1613 cm⁻¹, bottom) and transient (1635 cm⁻¹, top) ns-TRIR bands of RubpyamideBPM in D2O solution.
Transient ν(CO) bands are seen at both lower and higher energy than the ground state at 1652 and 1623 cm\(^{-1}\) respectively, consistent with population of both \(^3\)MLCT(bpyam) and \(^3\)MLCT(bpm) excited states. This has been confirmed through scaled subtraction and fitting analysis (not shown). As in D\(_2\)O, the bleach appears to grow in with a time constant of 4 (±2) ps due to heavy overlap between bleach and transient bands leading to excited state bands ‘imprinting’ their dynamics on the single-pixel bleach trace. The lower-energy ν(CO) transient overlaps heavily with the bleach of the ground-state band, preventing the straightforward extraction of kinetic information. The ν(CO) transient at higher energy grows in with a time constant of 6 (±2) ps and there are no subsequent changes in spectral profile on the picosecond timescale. The band at 1580 cm\(^{-1}\) arises from a bpm ring mode is bleached and a lower energy transient grows in at 1559 cm\(^{-1}\), although the data was not of sufficient quality to extract kinetic information. These results indicate the observation of excited state interconversion from the higher lying \(^3\)MLCT(bpyam) to the lower lying \(^3\)MLCT(bpm) states on a timescale which is the same, within experimental error, of the rate observed in D\(_2\)O of 7 (±2) ps.

ns-TRIR spectroscopy of RubpyamideBPM following 355 nm excitation shows that the amide bleach (at 1641 cm\(^{-1}\)) and transient (at 1652 cm\(^{-1}\)) bands then decay at the same rate, with a time constant of 77 (±4.6) ns. The agreement of the two lifetimes obtained confirms similarity of the photophysical properties of RubpyamideBPM in the two solvents.
Figure 4.4.3.5: The FTIR (bottom) and ps-TRIR spectra between 1 and 1000 ps (top) of RubpyamideBPM in CH$_3$CN solution.

Figure 4.4.3.6: Single-pixel kinetics of bleach (1638 cm$^{-1}$, bottom) and transient (1652 cm$^{-1}$, top) ps-TRIR bands of RubpyamideBPM in CH$_3$CN solution.
Figure 4.4.3.7: The FTIR (bottom) and ns-TRIR spectra between 1 and 500 ns (top) of RubpyamideBPM in CH$_3$CN solution.

Figure 4.4.3.8: Single-pixel kinetics of bleach (1637 cm$^{-1}$, bottom) and transient (1652 cm$^{-1}$, top) ns-TRIR bands of RubpyamideBPM in CH$_3$CN solution.
4.4.4. RubpyamideBPM: Closing Points

The results obtained for RubpyamideBPM in D₂O and CH₃CN suggest that the lowest energy excited state is Ru → bpm ³MLCT in nature, although there is evidence for the formation of a short-lived Ru → bpyam ³MLCT excited state immediately following excitation on the picosecond timescale. This is consistent with our model which is based on Ru → bipyrimidine chromophores possessing lower ³MLCT energies than their Ru → bipyridine counterparts. A comparison of ReBPM, RuCNBPM and RubpyamideBPM with the Rubpyamide-BLₓ complexes presented in Chapter 3 suggests that the principal findings when comparing monometallic bpm and bpy complexes are the existence of additional lower energy MLCT chromophores, a more facile 1st reduction, similar FTIR profile and significantly shorter lifetime and lower quantum yield. The results obtained for D₂O and CH₃CN are closely comparable and reiterate the lack of solvent perturbation of the electronic structure of RubpyamideBPM.

4.5. {(OC)₃ClRe(μ-bpm)ReCl(CO)₃}, (Re)₂BPM

(Re)₂BPM is the homobimetallic analogue of ReBPM (§4.2), with a second fac-[ReCl(CO)₃] unit appended to the bpm ligand. There are a few studies of (Re)₂BPM reported in the literature (§4.1) using absorption and FTIR spectroscopy and electrochemistry. We have repeated these experiments here and have conducted IR OTTLE and ps-TRIR experiments in order to understand the excited state dynamics of this system. (Re)₂BPM is neutral and non-polar, and as such is insoluble in
aqueous media. As the complex is only slightly solvatochromic and as its properties change little with solvent the results presented in CH$_3$CN are taken to be representative of other solvents. Only one of the two possible configurational isomers has been drawn above – it is uncertain whether the chlorines are syn or anti to each other but this is not expected to have measurable influence on the photophysics of the system (§4.1.2).

4.5.1. UV/visible Absorption and Luminescence Spectroscopy

The UV/visible absorption spectrum of (Re)$_2$BPM has been measured in CH$_3$CN (see Appendix 4.12 for the spectrum). The results are presented in Table 4.5.1.1 and agree well with previously published results (Table 4.1.2.2). The two absorption bands are shifted to significantly longer wavelengths ($\Delta \lambda = 50$–100 nm) than in ReBPM, confirming that complexation of a second metal centre decreases the HOMO-LUMO energy gap considerably. No luminescence was detected from (Re)$_2$BPM in CH$_3$CN solution, consistent with previous observations.

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<th>Solvent</th>
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<td>CH$_3$CN</td>
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<td>352</td>
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<tr>
<td></td>
<td>484</td>
</tr>
<tr>
<td>Assignment</td>
<td>IL $\pi$-$\pi^*$ (bpm)</td>
</tr>
</tbody>
</table>

*Table 4.5.1.1: Table of peak positions and assignments for UV/visible absorption spectra of (Re)$_2$BPM in CH$_3$CN solution.*

4.5.2. Cyclic Voltammetry

Cyclic voltammetry of (Re)$_2$BPM in CH$_3$CN was performed (see Appendix 4.12 for voltammograms) and the results are presented in Table 4.2.2.1 below. Features corresponding to two reductions (one reversible and one irreversible) and one irreversible oxidation couple are observed. The results are comparable to previous studies of this system.
<table>
<thead>
<tr>
<th>$E_{V}/V$ (vs. Fe/Fc$^+$)</th>
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<th>Assignment</th>
</tr>
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<tr>
<td>$E_p = +1.30$</td>
<td>Irreversible</td>
<td>Ru$^{III}$</td>
</tr>
<tr>
<td>-0.84</td>
<td>Reversible</td>
<td>(bpm)$^{0/+}$</td>
</tr>
<tr>
<td>-1.52</td>
<td>Reversible</td>
<td>(bpm)$^{+/2-}$</td>
</tr>
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</table>

Table 4.5.2.1: Table of potentials for redox couples and assignments from the cyclic voltammogram of (Re)$_2$BPM in CH$_3$CN solution.

4.5.3. FTIR, IR OTTLE and Time-Resolved Infrared Spectroscopy

(Re)$_2$BPM has been studied using FTIR, IR OTTLE and ps-TRIR spectroscopy in CH$_3$CN solution. Figure 4.5.3.1 shows the FTIR and ps-TRIR spectra obtained (Re)$_2$BPM in CH$_3$CN solution. Figure 4.5.3.2 compares the IR OTTLE (1$^{st}$ reduction) and TRIR difference spectra obtained in CH$_3$CN solution. Figure 4.5.3.3 shows the ps-TRIR kinetics obtained for bleach and transient bands in CH$_3$CN.

Acetonitrile Solution

FTIR Spectroscopy

The FTIR spectrum of (Re)$_2$BPM in CH$_3$CN solution in the $\nu$(CO) region has been measured. In the carbonyl region there is a peak profile characteristic of the rhenium (I) tricarbonyl unit with 3 resolved maxima at 2029, 1934 and 1921 cm$^{-1}$. This is in good agreement with previously published studies.

Infrared Spectroelectrochemistry

The FTIR spectrum of (Re)$_2$BPM in CH$_3$CN was monitored as the complex underwent its 1$^{st}$ reduction. The bands originating from the neutral complex at 2029, 1934 and 1921 cm$^{-1}$ decrease in intensity and newly formed bands centred at 2013, 1912 and 1896 cm$^{-1}$ become evident. These bands at lower wavenumber correspond to an increase in electron density at the rhenium centres, consistent with the reduced electron being located on the bridging bipyrimidine ligand.

4-32
ps-TRIR Spectroscopy

The ps-TRIR spectral profile of (Re)\textsubscript{2}BPM measured in CH\textsubscript{3}CN following 400 nm excitation contains three bleaches at 2029, 1934 and 1921 cm\textsuperscript{-1}, and two sets of three transient bands are detected at lower and higher energy to the bleaches respectively. The bands to higher energy are located at 2090, 1998 and 1972 cm\textsuperscript{-1} and correspond to a decrease in electron density at the metal centre as a result of a Re \rightarrow bpm \textsuperscript{3}MLCT transition taking place. The bands to lower energy are centred at 2015, 1900 and 1887 cm\textsuperscript{-1} and show good correlation with the IR bands of the reduced species [Re(bpm\textsuperscript{=} )Re] formed in the IR OTTLE experiment described above and shown in Figure 4.5.3.2. These bands are assigned to one ReCO terminus experiencing the increase in electron density at the bridging bipyrimidine (the 'spectator') as a result excitation of the other terminus. The \textsuperscript{3}MLCT excited state is best formulated as having an asymmetric charge distribution, such that the transient structure may be written in the form [Re(bpm\textsuperscript{=} )Re\textsuperscript{+}]. These results are consistent with excited state localisation on the IR timescale (§1.3), similar to other bimetallic rhenium (I) complexes previously studied using ps-TRIR spectroscopy (§4.1.3).\textsuperscript{10}
Figure 4.5.3.2: A comparison of TRIR spectrum at 100 ps (top), IR OTTLE difference spectrum (middle) and FTIR spectrum (bottom) of (Re)₂BPM in CH₃CN.

Figure 4.5.3.3: Single-pixel kinetics of bleach (2026 cm⁻¹, bottom) and transient (2013 cm⁻¹, top) ps-TRIR bands of (Re)₂BPM in CH₃CN solution.
A monoexponential decay is observed following vibrational cooling, with $\tau = 46 (\pm 1)$ ps (averaged over both bleach and transient features). This is significantly shorter than the excited state lifetime of ReBPM, confirming that complexation of a second metal centre reduces the HOMO-LUMO energy gap further, leading to a decrease in excited state lifetime. This is consistent with the 'energy-gap law' and the lack of any detected emission from this complex.

### 4.5.4. (Re)$_2$BPM: Closing Points

The lowest-lying excited state of (Re)$_2$BPM has been confirmed to be Re $\rightarrow$ bpm $^3$MLCT with the ps-TRIR data suggesting that the excited state is localised on the IR timescale (§1.3), in good agreement with ps-TRIR results reported for similar complexes (§4.1.3).$^{10}$ The complex exhibits longer wavelength absorption bands and more facile $1^{st}$ reduction than its model complex ReBPM, indicating that the complexation of a second metal to bpm reduces the HOMO-LUMO energy gap significantly through lowering of the bpm LUMO energy.

No other excited states have been observed which is significant as previous studies attempted to rationalise the lack of luminescence from this complex in terms of an unidentified low energy "dark" excited state. The TRIR data presented here does not provide any evidence for such a state, although the "snapshot-in-time" afforded by pulsed laser experiments does not provide information on the timescale of excited state localisation/delocalisation beyond the IR timescale (§1.3). In this case a delocalised-to-localised transition could be a possible explanation for the identity of the low energy absorption bands (and a "dark" excited state), but no evidence for further absorption bands in the nIR region has been observed. Such a band would be strong evidence for charge transfer transitions of a different nature e.g. MMCT or IVCT.
4.6. \( X_4\{(\text{NC})_4\text{Ru(\(\mu\)-bpm)}\text{Ru(CN)}_4\} \) (\( X = \text{K, PPN} \)), \((\text{RuCN})_2\text{BPM}\)

\((\text{RuCN})_2\text{BPM}\) is the homobimetallic analogue of \(\text{RuCNBPM}\) (§4.3), with a second \([\text{Ru(CN)}_4]\) unit complexed to the bpm ligand. Unlike \((\text{Re})_2\text{BPM}\), \((\text{RuCN})_2\text{BPM}\) is highly charged, strongly solvatochromic (by virtue of its 8 cyanide ligands) and polar. It has been characterised in \(\text{D}_2\text{O}\) and \(\text{CH}_3\text{CN}\) and by analogy with other ruthenium (II) cyanide complexes is expected to display markedly different spectroscopic, photophysical and electrochemical properties in the two solvents.

### 4.6.1. UV/visible Absorption and Luminescence Spectroscopy

The UV/visible absorption spectra of \((\text{RuCN})_2\text{BPM}\) has been measured in \(\text{CH}_3\text{CN}\) and \(\text{D}_2\text{O}\) (see Appendix 4.12 for spectra). The strong negative solvatochromism is clearly evident in the differences in peak positions in the two solvents. The results are summarised in Table 4.6.1.1. No luminescence from \((\text{RuCN})_2\text{BPM}\) was detected in either solvent.

<table>
<thead>
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<th>Solvent</th>
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</tr>
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<tbody>
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<td>(\text{D}_2\text{O})</td>
<td>&lt; 300 362 492 524</td>
</tr>
<tr>
<td>(\text{CH}_3\text{CN})</td>
<td>&lt; 300 400 434 478</td>
</tr>
<tr>
<td>Tentative Assignment</td>
<td>IL ( \pi-\pi^* ) (bpm) 1MLCT (Ru) 1MLCT (Ru) 1MLCT</td>
</tr>
</tbody>
</table>

*Table 4.6.1.1: Table of peak positions and assignments for UV/visible absorption spectra of \((\text{RuCN})_2\text{BPM}\) in \(\text{D}_2\text{O}\) and \(\text{CH}_3\text{CN}\) solutions.*
4.6.2. Cyclic Voltammetry

Cyclic voltammetry of \((\text{RuCN})_2\text{BPM}\) was measured in CH\(_3\)CN (see Appendix 4.12 for voltammograms) and the results are presented in Table 4.6.2.1 below. Features corresponding to one irreversible reduction and one irreversible oxidation couple are observed.

<table>
<thead>
<tr>
<th>(E_{\text{p}}/\text{V (vs. Fc/Fc\textsuperscript{+})})</th>
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<th>Assignment</th>
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<td>(E_{\text{p}} = -0.43)</td>
<td>Irreversible</td>
<td>(bpm)(0/\text{a-})</td>
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</table>

*Table 4.6.2.1: Table of potentials for redox couples and assignments from the cyclic voltammogram of \((\text{RuCN})_2\text{BPM}\) in CH\(_3\)CN solution.*

4.6.3. FTIR, IR OTTLE and Time-Resolved Infrared Spectroscopy

\((\text{RuCN})_2\text{BPM}\) has been studied using FTIR and ps-TRIR spectroscopy in D\(_2\)O and CH\(_3\)CN solution. IR OTTLE of \((\text{RuCN})_2\text{BPM}\) in CH\(_3\)CN was unsuccessful as both oxidation and reduction couples investigated were found to be chemically irreversible on the timescale of the experiment.

Deuterium Oxide Solution

Figure 4.6.3.1 shows the FTIR and ps-TRIR spectra of \((\text{RuCN})_2\text{BPM}\) measured in D\(_2\)O solution. Figure 4.6.3.2 shows the ps-TRIR kinetics obtained for bleach and transient bands in D\(_2\)O.
Figure 4.6.3.1: The FTIR (bottom) and ps-TRIR spectra between 1 and 1000 ps (top) of (RuCN)_2BPM in D_2O solution.

Figure 4.6.3.2: Single-pixel kinetics of bleach (2056 cm\(^{-1}\), bottom) and transient (2022 cm\(^{-1}\) top) ps-TRIR bands of (RuCN)_2BPM in D_2O solution.
The FTIR spectrum of (RuCN)_2BPM in D_2O solution in the ν(CN) region has a peak profile characteristic of the tetracyanoruthenate (II) unit with maxima at 2096, 2063 and 2046 cm\(^{-1}\). The ps-TRIR spectral profile consists of two resolved bleaches centred at 2092 and 2056 cm\(^{-1}\) and transient bands at lower and higher wavenumber respectively. The weak band at higher wavenumber overlaps with the small high energy bleach and is centred at ca. 2094 cm\(^{-1}\). This is consistent with a decrease in electron density as a result of Ru → bpm \(^3\)MLCT excited state formation. The transient band to lower wavenumber is considerably more intense than the other transient and is centred at 2024 cm\(^{-1}\). As in (Re)_2BPM this band is associated with a 'spectating' metal centre experiencing a net increase in electron density due to the MLCT transition occurring on the adjacent metal centre, and is thus indicative of asymmetric charge distribution in the excited state resulting from a localised excited state viz. [RuCN(bpm\(^-\))RuCN\(^+\)].

A monoexponential decay is observed with \(τ = 1200 (±85)\) ps (averaged over both bleach and transient features). This is significantly shorter then the excited state lifetime of RuCNBPM as expected for a smaller HOMO-LUMO energy gap after complexation of the second [Ru(CN)\(_4\)]\(^2-\) unit.\(^{15}\)

**Acetonitrile Solution**

Figure 4.6.3.3 shows the FTIR and ps-TRIR spectra obtained of (RuCN)_2BPM in CH\(_3\)CN solution. Figure 4.6.3.4 shows the ps-TRIR kinetics obtained for bleach and transient bands in CH\(_3\)CN. The FTIR spectrum of (RuCN)_2BPM in CH\(_3\)CN solution in the ν(CN) region has 3 resolved maxima at 2088, 2055 and 2038 cm\(^{-1}\). The profile obtained following 400 nm excitation consists of bleaches centred at 2088, 2055 and 2038 cm\(^{-1}\) and transient bands at lower and higher wavenumber, centred at 2015 and 2067 cm\(^{-1}\) respectively. The relative intensities of these bands receive close scrutiny below as the transient at 2067 cm\(^{-1}\) is barely detectable (presumably due to heavy overlap with bleach bands) whilst the transient band at 2015 cm\(^{-1}\) dominates the spectrum, having a considerably larger absorbance and width than the sum total of the three bleach bands.

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Figure 4.6.3.3: The FTIR (bottom) and ps-TRIR spectra between 1 and 1000 ps (top) of (RuCN)$_2$BPM in CH$_3$CN solution.

Figure 4.6.3.4: Single-pixel kinetics of bleach (2055 cm$^{-1}$, bottom) and transient (20/3 cm$^{-1}$, top) ps-TRIR bands of (RuCN)$_2$BPM in CH$_3$CN solution.
A monoexponential decay is observed in CH$_3$CN with $\tau = 65 (\pm 2)$ ps (averaged over both bleach and transient features). This is significantly shorter than the excited state lifetime of (RuCN)$_2$BPM in D$_2$O, which can be attributed to two principal reasons: the attenuation of decay pathways in deuterated solvents (§2.1.2),$^{17}$ and the ‘energy-gap law’.$^{15}$

The lower energy transient band associated with the ‘spectating’ [Ru$^{II}$CN$_4$]$^{2-}$ terminus has (in both CH$_3$CN and D$_2$O) a much higher intensity than the weak high energy transient band associated with the {Ru$^{III}$CN$_4$}$^-$ terminus and the relative intensities of the transients vary substantially between the two solvents. A rationalisation may be attempted on the basis of the change in the dipole moment of the C≡N bonds following charge redistribution in the excited state.$^{18}$ The increase in electron density at the spectating Ru centre in the excited state leads to an increase in the negative charge on the C atoms and hence an increase in the effective dipole moment of the C≡N oscillators, giving rise to more intense v(CN) bands. The converse is true for the transiently oxidised Ru centre, whose v(CN) band in the excited state is correspondingly weak because transient oxidation of the metal centre reduces the partial negative charge on C and hence diminishes the C-N$^+$ dipole, decreasing the IR oscillator strength.

There is a caveat emptor to consider here; the rationalisation above on IR intensity rationalisations is based on dipole (i.e. 1$^{st}$ order) arguments and there is some debate in the scientific literature as to the significance of quadrupolar (or even octupolar) interactions, particularly in metal cyanide complexes.$^{18}$ Our data don’t provide any insight into this phenomenon but it is interesting subject matter for a future study. Given that the author’s first published paper detailed the detection of an unexpectedly stable dipole-octupole complex in low-temperature matrices using FTIR spectroscopy$^{19}$ such “minor” effects should never be underestimated!
4.6.4. (RuCN)\(_2\)BPM: Closing Points

The lowest-lying excited state of (RuCN)\(_2\)BPM has been confirmed to be Ru \(\rightarrow\) bpm \(^3\)MLCT with the ps-TRIR data suggesting that the excited state is localised on the IR timescale (§1.3). The complex exhibits longer wavelength absorption bands than its model complex RuCNBPM, reinforcing the assertion that the complexation of a second metal to bpm considerably reduces the HOMO-LUMO energy gap through lowering of the bpm LUMO energy. The system exhibits markedly different TRIR spectroscopy and kinetics in the two solvent environments used, consistent with SSDA interactions between the solvent sphere and the system’s eight cyanide ligands.

4.7. \(((5,5'-\text{CONEt}_2)_2\text{bpy})_2\text{Ru(μ-bpm)}\text{Ru}((5,5'-\text{CONEt}_2)_2\text{bpy})_2\)Cl\(_4\), (Rubpyamide)\(_2\)BPM

(Rubpyamide)\(_2\)BPM is the homobimetallic analogue of RubpyamideBPM (§4.4), with a second [Ru((5,5'-CONEt\(_2\))\(_2\)bpy)]\(^4+\) unit appended to the bpm ligand. Complexes similar to (Rubpyamide)\(_2\)BPM have been studied before in the literature, most notably [(bpy)\(_2\)Ru(μ-bpm)Ru(bpy)]\(^4+\) (§4.1). Comparison of the results with those of RubpyamideBPM will provide insight into the ground and excited state electronic structure of these complexes.
(Rubpyamide)$_2$BPM is highly charged and polar, and as such is soluble in a wide range of solvents. The results obtained in D$_2$O and CH$_3$CN will be discussed here. As the complex is only slightly solvatochromic, small differences in properties are expected between the two solvents.

4.7.1. UV/visible Absorption and Luminescence Spectroscopy

The UV/visible absorption spectra of (Rubpyamide)$_2$BPM have been measured in D$_2$O and CH$_3$CN (see Appendix 4.12 for spectra) and the results are summarised in Table 4.7.1.1. The absorption spectra of these complexes are only slightly affected by solvent, with differences of a few nanometers wavelength between related absorption maxima in D$_2$O and CH$_3$CN. As with RubpyamideBPM more low energy absorption bands are observed in comparison with the other homobimetallic complexes, owing to the presence of Ru → bpyamide $^1$MLCT transitions in addition to the Ru → bpm $^1$MLCT transitions. There are two transitions, at ca. 420 and 400 nm, whose positions do not vary significantly with either solvent or nuclearity (when analysed in conjunction with RubpyamideBPM) and hence are assigned to the Ru → bpyamide $^1$MLCT transitions. The bands at ca. 530 and 360 nm are assigned to two Ru → bpm $^1$MLCT transitions, consistent with the presence of two different low-lying $\pi^*$ bpm orbitals giving two $^1$MLCT transitions to the bpm ligand. These results correlate well with a previous study of [(Ru(bpY)$_2$)(ll-bpm)]$^{4+}$.7

<table>
<thead>
<tr>
<th>Solvent</th>
<th>$\lambda_{\max}$/ nm</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>D$_2$O</td>
</tr>
<tr>
<td></td>
<td>&lt; 300</td>
</tr>
<tr>
<td></td>
<td></td>
</tr>
<tr>
<td>Assignment</td>
<td></td>
</tr>
<tr>
<td>IL $\pi$-$\pi^*$ (bpm)</td>
<td>$^1$MLCT (Ru-bpm)</td>
</tr>
</tbody>
</table>

Table 4.7.1.1: Table of peak positions and assignments for UV/visible absorption spectra of (Rubpyamide)$_2$BPM in D$_2$O and CH$_3$CN solution.

The luminescence of (Rubpyamide)$_2$BPM in CH$_3$CN is much weaker than that of RubpyamideBPM but detectable at room temperature, with a broad structureless
emission peak centred at 774 nm (see Appendix 4.12 for spectra). The luminescence intensity was insufficient to determine the excited state lifetime or quantum yield.

4.7.2. Cyclic Voltammetry

Cyclic voltammetry of (Rubpyamide)$_2$BPM was performed in CH$_3$CN (see Appendix 4.12 for voltammograms) and the results are presented in Table 4.7.2.1 below. Features corresponding to four reversible reductions and one irreversible oxidation couple are observed. The results are comparable to those obtained for RubpyamideBPM, with the notable difference of a significantly lower 1$^{st}$ reduction potential. This is consistent with a lowering of the LUMO energy of the bpm ligand upon complexation of the second metal centre.

<table>
<thead>
<tr>
<th>$E_{p}$/V (vs. Fe/Fe$^+$)</th>
<th>Reversibility</th>
<th>Assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>$E_p = +1.54$</td>
<td>Irreversible</td>
<td>Amide oxidation</td>
</tr>
<tr>
<td>-0.68</td>
<td>Reversible</td>
<td>(bpm)$^0$-</td>
</tr>
<tr>
<td>-1.35</td>
<td>Reversible</td>
<td>Amide reduction</td>
</tr>
<tr>
<td>-1.65</td>
<td>Reversible</td>
<td>Amide reduction</td>
</tr>
<tr>
<td>-1.85</td>
<td>Reversible</td>
<td>Amide reduction</td>
</tr>
</tbody>
</table>

*Table 4.7.2.1: Table of potentials for redox couples and assignments from the cyclic voltammogram of (Rubpyamide)$_2$BPM in CH$_3$CN solution.*

4.7.3. FTIR, IR OTTLE and Time-Resolved Infrared Spectroscopy

(Rubpyamide)$_2$BPM has been studied using FTIR and ps and ns-TRIR spectroscopy in D$_2$O and CH$_3$CN solution. An IR OTTLE experiment to monitor the changes in the FTIR spectrum of the complex in CH$_3$CN as it underwent its 1$^{st}$ reduction showed no spectral changes. As was the case with RubpyamideBPM this is consistent with little electronic coupling between the bipyrimidine bridging ligand and the bpyamide ligands.
Deuterium Oxide Solution

Figures 4.7.3.1 and 4.7.3.3 show the FTIR, ps-TRIR and ns-TRIR spectra obtained of (Rubpyamide)₂BPM in D₂O solution. Figures 4.7.3.2 and 4.7.3.4 show the ps and ns-TRIR kinetics obtained for bleach and transient bands in D₂O. The FTIR spectrum of (Rubpyamide)₂BPM in D₂O solution in the amide v(CO) region has a single amide peak present at 1614 cm⁻¹.

Picosecond TRIR spectra show that upon excitation the parent v(CO) band is bleached and a transient band is observed at higher energy centred at 1634 cm⁻¹. The bleach band appears to grow in with a time constant of 7 (±1) ps. This is thought to be an artefact due to heavy overlap with associated transient bands (see below) which imprint their dynamics on the single-pixel bleach trace as observed in other complexes containing the bpyamide ligand in this Chapter. As with RubpyamideBPM initial excitation is non-selective so both [(bpyam)(bpyam⁺)Ru³⁺(bpm)Ru²⁺(bpyam)] (hereafter ³MLCT(bpyamide)) and [(bpyam)₂Ru³⁺(bpm⁺)Ru²⁺(bpyam)] (hereafter ³MLCT(bpm)) triplet excited states are expected on the picosecond timescale, the latter being significantly lower in energy (as demonstrated by UV/vis spectroscopy, vide supra). As with RubpyamideBPM, the transient v(CO) band which is shifted to higher energy can be used to monitor the formation and decay of the ³MLCT(bpm) excited state. In D₂O this transient grows in with a time constant of 9 (±1) ps. In contrast to the results obtained for RubpyamideBPM, no transient band to lower wavenumber is observed. This has been confirmed through scaled subtraction and fitting analysis (not shown). There are no further significant changes during the picosecond timescale.

ns-TRIR spectroscopy (with excitation at 355 nm) gives an analogous profile to the ps-TRIR spectra (vide supra), with the parent bleach centred at 1614 cm⁻¹ and a transient centred at 1634 cm⁻¹. These bands decay synchronously with a lifetime of 6.8 (±1.6) ns. This behaviour is consistent with internal conversion of the higher-lying ³MLCT(bpyamide) state to the lower-lying ³MLCT(bpm) state, formally by internal
Figure 4.7.3.1: The FTIR (bottom) and ps-TRIR spectra between 1 and 1000 ps (top) of (Rubpyamide)\textsubscript{2}BPM in D\textsubscript{2}O solution.

Figure 4.7.3.2: Single-pixel kinetics of bleach (1611 cm\textsuperscript{-1}, bottom) and transient (1635 cm\textsuperscript{-1}, top) ps-TRIR bands of (Rubpyamide)\textsubscript{2}BPM in D\textsubscript{2}O solution.
Figure 4.7.3.3: The FTIR (bottom) and ns-TRIR spectra between 1 and 100 ns (top) of (Rubpyamide)₂BPM in D₂O solution.

Figure 4.7.3.4: Single-pixel kinetics of bleach (1613 cm⁻¹, bottom) and transient (1637 cm⁻¹, top) ns-TRIR bands of (Rubpyamide)₂BPM in D₂O solution.
electron transfer from (bpyamide)$^-$ to bpm, followed by slower decay of the $^3$MLCT(bpm) state to the ground state on the nanosecond timescale.

**Acetonitrile Solution**

Figures 4.7.3.5 and 4.7.3.7 show the ps and ns-TRIR spectra of RubpyamideBPM taken in CH$_3$CN solution. Figures 4.7.3.6 and 4.7.3.8 show the ps and ns-TRIR kinetics obtained for bleach and transient bands in CH$_3$CN. The FTIR spectrum of (Rubpyamide)$_2$BPM in CH$_3$CN solution in the amide $\nu$(CO) region has a single amide peak present at 1635 cm$^{-1}$.

ps-TRIR spectroscopy of (Rubpyamide)$_2$BPM gives a similar bleach of the main $\nu$(CO) amide band following 400 nm excitation as in D$_2$O, but interestingly transient bands to lower and higher wavenumber are observed, centred at 1616 and 1646 cm$^{-1}$ respectively. As in D$_2$O, the bleach appears to grow in with a time constant of 4 (± 2) ps due to heavy overlap with transient bands imprinting their dynamics on the single-pixel kinetic trace. Furthermore, scaled subtraction and fitting analysis showed that this peak to lower wavenumber is an artefact of the very broad transient at 1646 cm$^{-1}$ compounded by the heavy overlap of bleach and transient bands. The $^3$MLCT(bpm) excited state then decays with a time constant of 6.2 (±0.7) ns.

The presence of the band at lower wavenumber is indicative of the detection of the $^3$MLCT(bpyamide) excited state, which undergoes rapid interconversion into the $^3$MLCT(bpm) state which is monitored through the band at 1616 cm$^{-1}$. Ostensibly this process is complete within 1 ps in D$_2$O, whereas in CH$_3$CN TRIR spectroscopy gives a time constant for $^3$MLCT(bpyamide) $\rightarrow$ $^3$MLCT(bpm) interconversion of 14 (±3) ps.
Figure 4.7.3.5: The FTIR (bottom) and ps-TRIR spectra between 1 and 1000 ps (top) of (Rubpyamide)₂BPM in CH₃CN solution.

Figure 4.7.3.6: Single-pixel kinetics of bleach (1635 cm⁻¹, bottom) and transient (1648 cm⁻¹, top) ps-TRIR bands of (Rubpyamide)₂BPM in CH₃CN solution.
Figure 4.7.3.7: The FTIR (bottom) and ns-TRIR spectra between 1 and 100 ns (top) of (Rubpyamide)$_2$BPM in CH$_3$CN solution.

Figure 4.7.3.8: Single-pixel kinetics of bleach (1636 cm$^{-1}$, bottom) and transient (1648 cm$^{-1}$, top) ns-TRIR bands of (Rubpyamide)$_2$BPM in CH$_3$CN solution.
4.7.4. (Rubpyamide)$_2$BPM: Closing Points

The lowest-lying excited state of (Rubpyamide)$_2$BPM has been confirmed to be Ru $\rightarrow$ bpm $^3$MLCT with the TRIR data ostensibly suggesting that the excited state is localised on the IR timescale (§1.3). The initial formation and subsequent decay of a higher energy Ru $\rightarrow$ bpyamide $^3$MLCT excited state has been observed in some cases. The complex exhibits longer wavelength absorption bands and more facile 1$^{st}$ reduction than its model complex RubpyamideBPM, reiterating that complexation of a second metal to bpm considerably reduces the HOMO-LUMO energy gap through lowering of the bpm LUMO energy.

4.8. $X_2\{(O\text{C})_2\text{ReCl(}$µ-bpm$)$\text{Ru(CN)}\}_4$ (X = K, PPN), ReBPMRuCN

ReBPMRuCN is a heterobimetallic complex with ReCO and RuCN termini on either side of the bpm bridging ligand. Each terminus has an M $\rightarrow$ bpm MLCT chromophore associated with it. As the bipyrimidine bridge is conjugated and strong intermetallic electronic coupling has been demonstrated for the homobimetallic complexes studied earlier in this Chapter (vide supra) the ground and excited state properties of ReBPMRuCN are not necessarily a superposition of those observed for ReBPMRe and RuCNBPMRuCN.

Whilst the RuCN terminus is strongly solvatochromic, the ReCO terminus is essentially solvent-independent. This will ostensibly lead to differences in the photophysical and spectroscopic properties of ReBPMRuCN in D$_2$O and CH$_3$CN,

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although the differences in \(^1\)MLCT transition energies of ReBPM and RuCNBPBM (\(\Delta \text{ } ^1\text{MLCT} = 3300-9150 \text{ cm}^{-1}\) depending on solvent; calculated from absorption spectra) are of sufficient magnitude to render solvent-based switching of the location of the excited state unlikely; Ru \(\rightarrow\) bpm \(^3\)MLCT is expected to be the lowest energy excited state in both CH\(_3\)CN and D\(_2\)O.

4.8.1. UV/visible Absorption and Luminescence Spectroscopy

The UV/visible absorption spectra of ReBPMRuCN have been measured in D\(_2\)O and CH\(_3\)CN (see Appendix 4.12 for spectra). It is difficult to discern peak maxima in the spectra as very broad features dominate the spectrum across the visible region. The peaks maxima extracted are reported in Table 4.8.1.1. The emission spectrum of ReBPMRuCN has been measured in CH\(_3\)CN solution. No luminescence was detected.

<table>
<thead>
<tr>
<th>Solvent</th>
<th>(\lambda_{\text{max}} / \text{nm})</th>
</tr>
</thead>
<tbody>
<tr>
<td>D(_2)O</td>
<td>&lt;300 384 480</td>
</tr>
<tr>
<td>CH(_3)CN</td>
<td>&lt;300 328 434</td>
</tr>
<tr>
<td>Tentative Assignment</td>
<td>IL (\pi-\pi^*)  (^1)MLCT  (^1)MLCT</td>
</tr>
<tr>
<td>Location</td>
<td>bpm ReCO RuCN</td>
</tr>
</tbody>
</table>

*Table 4.8.1.1: Table of peak positions and assignments for UV/visible absorption spectra of ReBPMRuCN in D\(_2\)O and CH\(_3\)CN solutions.*

4.8.2. Cyclic Voltammetry

Cyclic voltammetry of ReBPMRuCN in D\(_2\)O has been measured (see Appendix 4.12 for voltammograms) and the results are presented in Table 4.8.2.1 below. Features corresponding to one reversible reduction and two oxidation couples (one reversible and one irreversible) are observed. It is noteworthy to mention the ease of reduction of the bipyrimidine ligand in this case – the redox couple’s proximity to 0 V indicates that this redox process may be accessible in the absence of electrical current, which
could have important implications for the electrochemical equilibrium mixture of electronic structures of the ground and excited states of ReBPMRuCN in solution.

<table>
<thead>
<tr>
<th>( E_{\text{p}} / \text{V} ) (vs. Fe/Fe(^+))</th>
<th>Reversibility</th>
<th>Tentative Assignment</th>
</tr>
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<td>( E_{\text{p}} = +1.17 )</td>
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<td>Re(^{\text{IV}})</td>
</tr>
<tr>
<td>(+0.87)</td>
<td>Reversible</td>
<td>Ru(^{\text{III}})</td>
</tr>
<tr>
<td>( E_{\text{p}} = +0.23 )</td>
<td>Reversible</td>
<td>((\text{bpm})^{0/-})</td>
</tr>
</tbody>
</table>

_Table 4.8.2.1: Table of potentials for redox couples and assignments from the cyclic voltammogram of ReBPMRuCN in CH\(_3\)CN solution._

4.8.3. FTIR, IR OTTLE and Time-Resolved Infrared Spectroscopy

ReBPMRuCN has been studied using FTIR and ps and ns-TRIR spectroscopy in D\(_2\)O and CH\(_3\)CN solutions. As a result of the negative solvatochromism of the RuCN terminus differences in the two solvents are expected, but the identity of the lowest energy excited is predicted to be Ru \(\rightarrow\) bpm \(^3\)MLCT is either case.

IR OTTLE of ReBPMRuCN in CH\(_3\)CN was unsuccessful as the 1\(^{st}\) reduction couple was found to be chemically irreversible on the timescale of the experiment.

**Deuterium Oxide Solution**

The FTIR spectrum of ReBPMRuCN has been measured in D\(_2\)O solution (Figure 4.8.3.1). In the cyanide region there is a peak profile characteristic of the tetracyanoruthenate (II) unit with resolved maxima at 2071 and 2045 cm\(^{-1}\). In the carbonyl region there are peaks present at 2022 and 1908 cm\(^{-1}\). These results are broadly comparable with the FTIR spectra of the model complexes ReBPM and RuCNBPM (§4.2.3, §4.3.3).
Picosecond and nanosecond TRIR spectroscopy (with laser excitation at 400 and 355 nm respectively) were performed on ReBPMRuCN in D$_2$O solution, monitoring both regions in the IR where reporter groups are present.

**ps-TRIR Spectroscopy**

Figure 4.8.3.1 shows the ps-TRIR spectra obtained for the ReCO and RuCN termini in D$_2$O. Figure 4.8.3.2 shows the ps-TRIR kinetics obtained for bleach and transient bands in D$_2$O. It is immediately apparent that the features arising from the two termini overlap, giving rise to a dense, feature-rich spectrum between 2100 and 1800 cm$^{-1}$. By analogy with (Re)$_2$BPMRe and (RuCN)$_2$BPM (§4.5.3, §4.6.3) we would expect TRIR spectroscopy to detect a localised excited state (on the IR timescale) with one set of transient bands shifted to higher energy than their bleach, and one set shifted lower. Combining this with the unselective (400 nm) excitation user for the ps-TRIR experiments, this would give bands at higher and lower wavenumber to the parent bleaches for both termini. This has important implications for the following analysis.

**RuCN (Lower Energy) terminus**

The parent bleach of the RuCN terminus is barely visible at 2051 cm$^{-1}$ as it is surrounded by a small transient to higher wavenumber centred at 2092 cm$^{-1}$ and a broad, intense transient to lower energy centred at about 1979 cm$^{-1}$, which dominates the spectrum. These two transient features can be assigned to Ru $\rightarrow$ bpm and Re $\rightarrow$ bpm $^3$MLCT excited states respectively. In the latter case the RuCN terminus is sensing the increase in electron density on the bpm bridging ligand as a result of a MLCT transition involving the other metal centre. There are strong similarities to the spectra of (RuCN)$_2$BPM and the reasons for the low energy feature’s intensity are explained in detail in §4.6.3.

The RuCN bands decay polyexponentially with a very fast component [$\tau_1 < 1$ ps] and a slower component $\tau_2 = 135$ (±15) ps. At $\Delta t = 1$ ns less than ca. 5% of the signal at $\Delta t = 1$ ps remains, indicating that quenching of the excited state is almost complete during the picosecond time domain.
ReCO (Higher Energy) terminus

The ReCO parent bleaches are apparent at 2011 and 1905 cm\(^{-1}\), with the former being detected as a sharp depletion in the large RuCN transient band which dominates the spectrum (vide supra). Transient bands are clearly evident at 1990 and 1854 cm\(^{-1}\), which are at lower wavenumber to the bleaches and correspond to the Re centre experiencing a net increase in electron density on the bridging ligand arising from Ru → bpm \(^3\)MLCT. There is the distinct possibility that transients to higher wavenumber than the bleaches may also be present but are obscured by the numerous features detected.

The ReCO bands decay polyexponentially with a very fast component (\(\tau_1 < 2\) ps) and a slower component with \(\tau_2 = 34\) (±10) ps (averaged over bands at 1993, 1901 and 1855 cm\(^{-1}\)). Due to the overlap of the bleach at 2011 cm\(^{-1}\) with RuCN features no meaningful kinetics could be obtained in that region. At \(\Delta t = 1\) ns there are barely detectable signals remaining. This adds further evidence to the earlier supposition that quenching of the excited state is almost complete during the picosecond time domain.
The ps-TRIR spectra recorded do not show any evidence for conversion between the Re → bpm and Ru → bpm $^3$MLCT excited states, but this process may be occurring on the sub-picosecond timescale.

**ns-TRIR Spectroscopy and Kinetics**

Attempts to measure the ns-TRIR spectrum of ReBPMRuCN in D$_2$O were unsuccessful, ostensibly due to insufficient signal-to-noise ratios obtained following the rapid excited state decay on the picosecond timescale.

**Acetonitrile Solution**

The FTIR spectrum of ReBPMRuCN in CH$_3$CN solution in the v(CN) and v(CO) regions has been measured. In the cyanide region there is a broad peak profile centred at 2067 cm$^{-1}$. In the carbonyl region there are peaks present at 2015 and 1899 cm$^{-1}$. These results show a degree of similarity to the FTIR spectra of the model complexes ReBPM and RuCNBPM (§4.2.3, §4.3.3). Picosecond and nanosecond TRIR spectroscopy (with laser excitation at 400 and 355 nm respectively) were performed on ReBPMRuCN in CH$_3$CN solution, monitoring both regions in the IR where reporter groups are present.

**ps-TRIR Spectroscopy**

Figure 4.8.3.4 shows the ps-TRIR spectra obtained for the ReCO and RuCN termini in CH$_3$CN and Figure 4.8.3.5 shows the ps-TRIR kinetics for bleach and transient bands in CH$_3$CN. The features arising from the two termini overlap, giving rise to a very feature-rich spectrum between 2100 and 1800 cm$^{-1}$, is a similar scenario to that observed in D$_2$O. Excited state localisation and unselective excitation should result in transient bands at higher and lower wavenumber to the parent bleaches for both termini. This has important implications for the following analysis.
Figure 4.8.3.2: ps-TRIR kinetics of ReBPMRuCN in D$_2$O solution for RuCN (transient at 2077 cm$^{-1}$, top left and bleach at 2056 cm$^{-1}$, bottom left) and ReCO (transient at 1993 cm$^{-1}$, top right and bleach at 1901 cm$^{-1}$, bottom right) termini for $\Delta t = 0$-1000 ps.
RuCN (Lower Energy) terminus

The parent bleach of the RuCN terminus is evident as a broad feature centred at 2058 cm\(^{-1}\) and a smaller depletion at 2100 cm\(^{-1}\). It is surrounded by a small transient to higher wavenumber centred at 2089 cm\(^{-1}\) and a broad, intense transient to lower energy centred at about 1985 cm\(^{-1}\) which dominates the spectrum. In a similar manner to spectra recorded in D\(_2\)O, these transient bands can be assigned to Ru \(\rightarrow\) bpm and Re \(\rightarrow\) bpm \(^2\)MLCT excited states respectively. In the latter case the RuCN terminus is sensing the increase in electron density on the bpm bridging ligand as a result of a MLCT transition involving the other metal centre. As in D\(_2\)O there are strong similarities to the spectra of (RuCN)\(_2\)BPM and the reasons for this feature’s intensity are explained in detail for in §4.6.3).

The RuCN bleach decays polyexponentially with a very fast component \((\tau_1 < 1\) ps) and a slower component \(\tau_2 = 135 \pm 15\) ps. At \(\Delta t = 1\) ns less than ca. 5% of the signal at \(\Delta t = 1\) ps remains, indicating that quenching of the excited state is almost complete during the picosecond time domain.

Figure 4.8.3.3: The FTIR (bottom) and ps-TRIR spectra between 1 and 1000 ps (top) of ReBPMRuCN in CH\(_3\)CN solution.
ReCO (Higher Energy) terminus

The ReCO parent bleaches are apparent at 2012 and 1901 cm\(^{-1}\), with the former being detected as a sharp depletion in the large RuCN transient band which dominates the spectrum (Figure 4.8.3.3). Transient bands are clearly evident at 1990 and 1855 cm\(^{-1}\), which are at lower wavenumber to the bleaches and correspond to the ReCO terminus experiencing a net increase in electron density on the bridging ligand arising from the formation of a Ru \(\rightarrow\) bpm \(^3\)MLCT excited state. It is conceivable that transients to higher wavenumber than the bleaches may also be present but may be obscured by transient features. The similarity of the peak positions in both solvents indicates that the electronic structures of the excited states are ostensibly similar.

The ReCO bands decay polyexponentially with a very fast component \((\tau_1 < 1\) ps) and a slower component with \(\tau_2 = 26 \pm 6\) ps (averaged over bands at 1990, 1893 and 1855 cm\(^{-1}\)). Due to the overlap of the bleach at 2012 cm\(^{-1}\) with RuCN features no meaningful kinetics could be obtained. At \(\Delta t = 1\) ns less than ca. 5% of the signal at \(\Delta t = 1\) ps remains. This adds further evidence to the earlier supposition that quenching of the excited state is almost complete during the picosecond time domain. As in D\(_2\)O there is no evidence for conversion of one state into another, but this process may be occurring on the femtosecond timescale.

Nanosecond Spectroscopy and Kinetics

Figures 4.8.3.6 and 4.8.3.7 show the ns-TRIR spectra and kinetics obtained in CH\(_3\)CN. As discussed above, most of the excited state bands decay within the picosecond timescale, so the ns-TRIR spectra probes the weak residual remaining features. The ps and ns-TRIR spectra exhibit major differences, the most striking of which is that the RuCN transient feature which dominates the ps-TRIR spectra is no longer apparent. On the RuCN terminus, the only features that remain are the parent bleach at 2068 cm\(^{-1}\) and its associated transient at 2094 cm\(^{-1}\), which corresponds to a Ru \(\rightarrow\) bpm \(^3\)MLCT excited state. On the ReCO terminus bleaches at 2012 and 1893 cm\(^{-1}\) and transient peaks at 1990 and 1924 cm\(^{-1}\) are observed. The results are consistent with one excited state persisting into the nanosecond timescale, in contrast
Figure 4.8.3.4: ps-TRIR kinetics of ReBPMRuCN in CH$_3$CN solution for RuCN (transient at 2083 cm$^{-1}$, top left and bleach at 2058 cm$^{-1}$, bottom left) and ReCO (transient at 1990 cm$^{-1}$, top right and bleach at 1893 cm$^{-1}$, bottom right) termini for $\Delta t = 0$-1000 ps.
Figure 4.8.3.5: The FTIR (bottom) and ns-TRIR spectra between 1 and 40 ns (top) of ReBPMRuCN in CH$_3$CN solution.

Figure 4.8.3.6: Single-pixel kinetics of bleach (2010 cm$^{-1}$, bottom) and transient (1995 cm$^{-1}$, top) ns-TRIR bands of ReBPMRuCN in CH$_3$CN solution.
to the two observed in the ps-TRIR spectra \((\text{vide supra})\). The ReCO bands decay with a time constant of \(\tau = 6.2 (\pm 0.4)\) ns (averaged over both bleach and transient values).

### 4.8.4. ReBPMRuCN: Closing Points

The lowest excited state of ReBPMRuCN is assigned to Ru → bpm \(^3\)MLCT, with transient bands from this state persisting into the nanosecond timescale in CH\(_3\)CN. The ps-TRIR spectra obtained in both solvents show evidence for excited states arising from both metal centres owing to unselective excitation, which contributes to the spectral complexity observed. There is no direct kinetic evidence for interconversion of excited states on the picosecond timescale; for example excited state interconversion of Re → bpm to Ru → bpm \(^3\)MLCT is energetically favourable in D\(_2\)O and CH\(_3\)CN. However such processes may feasibly be complete within 1 ps, particularly in conjugated systems.\(^{20}\) The spectral profiles observed are similar in both solvents, consistent with the expectation that excited state ordering is unaffected by changes in solvent Acceptor Number. The lack of luminescence is consistent with the rapid excited state deactivation of the majority of the TRIR signal. The polyexponential nature of the ps-TRIR kinetics is postulated to arise from the presence of additional \(^3\)MLCT states involved in the excited state relaxation process. These states are likely to possess similar energies to the lowest-lying excited Ru → bpm state due to the presence of two \(\pi^*\) LUMOs in the bpm ligand with similar energies (Figure 4.1.1.2). Another possible contributing factor is the high sample concentrations required experimentally to obtain good signal-to-noise ratios with the PIRATE apparatus, which may lead to self-quenching of excited state species (see Chapter 6 for more details).
4.9. \{(OC)₃ReCl(µ-bpm)Ru((5,5'-CONEt₂)₂bpy)₂\}₂Cl₂, ReBPMRubpyamide

ReBPMRubpyamide is a heterobimetallic complex with ReCO and Rubpyamide termini on either side of the bpm bridging ligand. Each terminus has an M → bpm MLCT chromophore associated with it, and the Rubpyamide terminus possesses an additional Ru → bpyamide chromophore. This may complicate the photophysical dynamics of the system in a similar manner to other Rubpyamide-containing complexes in this Chapter (§4.4 and §4.7) and could be an issue in the event of unselective excitation in the TRIR experiments. Neither terminus is strongly solvatochromic and as a result no significant differences in the spectroscopic properties of ReBPMRubpyamide are expected between D₂O and CH₃CN. Ru → bpm ³MLCT is expected to be the lowest energy excited state (and hence the longest-lived) in both CH₃CN and D₂O.

4.9.1. UV/visible Absorption and Luminescence Spectroscopy

UV/visible absorption spectra of ReBPMRubpyamide have been measured in D₂O and CH₃CN (see Appendix 4.12 for spectra). The results are presented in Table 4.9.1.1. The spectral profile obtained is similar to that reported for \[[\text{ReCl(CO)}₃(µ-bpm)(\text{Ru(bpy)}₂)]²⁺\] in CH₃CN.⁵
The luminescence spectrum of ReBPMRubpyamide was measured in CH$_3$CN solution (see Appendix 4.12 for spectra). The results are presented in Table 4.9.1.2. Excitation directly into the Ru → ((4,4'-CONEt$_2$)hbpy) MLCT chromophore at 426 nm results in weak luminescence with a broad, structureless profile centred at 646 nm, strongly resembling the emission observed in RubpyamideBPM. Excitation (at 550 nm) into the low energy charge transfer band gives barely detectable emission also centred at 646 nm. In both cases the emission intensity was insufficient to perform TCSPC experiments to determine the lifetime.

<table>
<thead>
<tr>
<th>$\lambda_{\text{ex}}$/ nm</th>
<th>$\lambda_{\text{em}}$/ nm</th>
<th>$\Phi$</th>
</tr>
</thead>
<tbody>
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<td>646</td>
<td>0.0006</td>
</tr>
<tr>
<td>550</td>
<td>646</td>
<td>0.0003</td>
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</tbody>
</table>

*Table 4.9.1.2: Table of peak positions and assignments for emission spectra of ReBPMRubpyamide in CH$_3$CN solution.*

### 4.9.2. Cyclic Voltammetry

Cyclic voltammetry of ReBPMRubpyamide was measured in D$_2$O (see Appendix 4.12 for the voltammogram) and the results are presented in Table 4.9.2.1 below. Features corresponding to three reversible reductions and three oxidation couples (one reversible and two irreversible) are observed. As with ReBpmRuCN, the ease of reduction of the bipyrimidine ligand may be significant. The redox couple’s proximity

<table>
<thead>
<tr>
<th>Solvent</th>
<th>$\lambda_{\text{max}}$/ nm</th>
</tr>
</thead>
<tbody>
<tr>
<td>D$_2$O</td>
<td>&lt; 300 323 358 394 418 535</td>
</tr>
<tr>
<td>CH$_3$CN</td>
<td>&lt; 300 325 355 401 426 544</td>
</tr>
</tbody>
</table>

*Table 4.9.1.1: Table of peak positions and assignments for UV/visible absorption spectra of ReBPMRubpyamide in D$_2$O and CH$_3$CN solutions.*
to 0 V indicates that this couple may be accessible in the absence of electrical current, which could have important implications for the electronic structure of the ground and excited states of ReBPMRubpyamide in solution. The voltammogram obtained is similar to that reported for [(ReCl(CO)3)(μ-bpm)(Ru(bpy)2)]2+ in CH3CN.5

<table>
<thead>
<tr>
<th>E½/V (vs. Fe/Fe+)</th>
<th>Reversibility</th>
<th>Tentative Assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>Eᵖ = +1.71</td>
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<td>Re⁷⁺</td>
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<tr>
<td>Eᵖ = +1.17</td>
<td>Irreversible</td>
<td>Ru⁷⁺</td>
</tr>
<tr>
<td>Eᵖ = +0.70</td>
<td>Irreversible</td>
<td>Amide Oxidation</td>
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<tr>
<td>-0.60</td>
<td>Reversible</td>
<td>(bpm)⁰⁻⁻</td>
</tr>
<tr>
<td>-1.37</td>
<td>Reversible</td>
<td>Amide Reduction</td>
</tr>
<tr>
<td>-1.65</td>
<td>Reversible</td>
<td>Amide Reduction</td>
</tr>
</tbody>
</table>

Table 4.9.2.1: Table of potentials for redox couples and assignments from the cyclic voltammogram of ReBPMRubpyamide in CH3CN solution.

4.9.3. FTIR, IR OTTLE and Time-Resolved Infrared Spectroscopy

ReBPMRubpyamide has been studied using ps and ns-TRIR spectroscopy in D₂O and CH₃CN solutions.

IR OTTLE of ReBPMRubpyamide in CH₃CN was unsuccessful as the 1ˢᵗ reduction couple was found to be chemically irreversible on the timescale of the experiment.

Deuterium Oxide Solution

The FTIR spectrum of ReBPMRubpyamide in D₂O solution in the ν(CO) and amide ν(CO) region has been measured. In the carbonyl region there is a peak profile characteristic of the rhenium (I) tricarbonyl unit with 3 resolved maxima at 2015, 1997 and 1925 cm⁻¹. There are also smaller ‘satellite’ peaks visible at 1874, 1835 and 1810 cm⁻¹. In the amide region there is a single peak present at 1616 cm⁻¹. These
results are broadly comparable with the FTIR spectra of the model complexes ReBPM and RubpyamideBPM (§4.2.3, §4.4.3), although ReBPM is insoluble in aqueous media and hence was only studied in CH$_3$CN. One significant difference from the model complexes’ spectra (also observed in the FTIR spectra of other bimetallics in this Chapter) is the presence of satellite peaks, at lower wavenumber to the carbonyl bands. This suggests that the 1st reduction of ReBPMRubpyamide ($E_{1/2} = -0.598$ V vs. Fc/Fc$^+$) is partially accessible in D$_2$O solution without any potential difference applied, and the peaks correspond to the FTIR spectrum of the partially reduced complex (with the electron residing on the bipyrimidine ligand). It is reasonable to suggest that the absence of such a satellite feature from the amide $v$(CO) region is due to the comparatively small amount of electronic coupling between the amide groups (situated on the bpy ligands) and the bipyrimidine bridge.

Picosecond and nanosecond-TRIR spectroscopy (with laser excitation at 400 and 355 nm respectively) were performed on ReBPMRubpyamide in D$_2$O solution, monitoring both regions in the IR where reporter groups are present.

**ps-TRIR Spectroscopy**

Figure 4.9.3.1 shows the ps-TRIR spectra obtained for the ReCO and Rubpyamide termini in D$_2$O. Figure 4.9.3.2 shows the ps-TRIR kinetics obtained for bleach and transient bands in D$_2$O.

**ReCO (“Donor” / Higher Energy) terminus**

The spectral profile obtained following 400 nm excitation is dominated by two bleaches at 2027 and 1936 cm$^{-1}$. Two transient bands are formed at lower energy to the bleaches, centred at 2015 and 1925 cm$^{-1}$. This is consistent with the ReCO terminus experiencing an increase in electron density as a result of a MLCT transition involving the Rubpyamide terminus (*vide infra*). There is considerable overlap between the bleaches and their associated transient bands. These bands narrow and blue-shift over the first 25 ps of the experiment, indicating that vibrational cooling takes place during this period. A biexponential decay is observed with $\tau_1 = 120 (\pm 75)$
and $\tau_2 = 1060 (\pm 195)$ ps (averaged over both bleach and transient features). At $\Delta t = 1$ ns only ca. 10% of the signal present at $\Delta t = 1$ ps remains, indicating that quenching of the excited state is almost complete during the picosecond time domain.

**Rubpyamide ("Acceptor" / Lower Energy) terminus**

The ps-TRIR spectra obtained are comprised of a parent bleach centred at 1616 cm$^{-1}$ and a transient band to higher wavenumber at 1636 cm$^{-1}$. This is consistent with a decrease in electron density on the bpy ligands, as expected during the formation of a Ru $\rightarrow$ bpm $^3$MLCT excited state. Additionally a small bleach is present at 1582 cm$^{-1}$ and is accompanied by a transient at 1560 cm$^{-1}$. By comparison with RubpyamideBPM and (Rubpyamide)$_2$BPM this can be assigned to a vibrational mode of the bipyrimidine ligand, directly reporting on the increase in electron density on the bridging ligand. Examination of the kinetics of the amide bands shows evidence for a grow-in of absorbance over the first 100 ps on both bleach and transient features. The duration of this growth and its manifestation in both bleach and transient spectra indicates that vibrational cooling is not responsible for this observation and there is a good case for it matching up with the fast decay process $\tau_1$ on the ReCO terminus.

After this grow-in has ended, the signal decays with a time constant (averaged over both bleach and transient features) of 1050 (±350) ps. This is in very good agreement with the slow decay process $\tau_2$ on the ReCO terminus (*vide supra*). At $\Delta t = 1$ ns ca. 25% of the signal at $\Delta t = 1$ ps remains, indicating that a long-lived component persists into the nanosecond time domain.

**ns-TRIR Spectroscopy**

Figures 4.9.3.3 and 4.9.3.4 show the ns-TRIR spectra and kinetics obtained respectively for the Rubpyamide terminus in D$_2$O. As with the ps-TRIR spectra, the spectrum is dominated by a bleach centred at 1616 cm$^{-1}$ with a transient band at 1633 cm$^{-1}$. A bleach band arising from a vibrational mode of bipyrimidine is detected at 1583 cm$^{-1}$ and its associated transient is centred at 1561 cm$^{-1}$. All the bands decay
synchronously with a biexponential time constants (averaged over both bleach and transient features) of $\tau_1 = 1.7 \pm 0.1$ and $\tau_2 = 63.2 \pm 2.5$ ns. The $\tau_1$ decay component from ns-TRIR spectroscopy shows relatively good agreement with $\tau_2$ in the picosecond spectra (vide supra), considering that neither the ps nor ns-TRIR experimental setups are well suited to measuring lifetimes of a few nanoseconds.

Acetonitrile Solution

The FTIR spectrum of ReBPMRubpyamide in CH$_3$CN solution in the v(CO) and amide v(CO) region has been measured. In the carbonyl region there is a peak profile characteristic of the rhenium (I) tricarbonyl unit with 3 resolved maxima at 2034, 1936 and 1919 cm$^{-1}$. There are also smaller ‘satellite’ peaks visible at 2017, 1903 and 1879 cm$^{-1}$. In the amide region there is a single peak present at 1639 cm$^{-1}$.

These results are comparable with the model complexes ReBPM and RubpyamideBPM in CH$_3$CN solution (§4.2.3, §4.4.3). As in D$_2$O, the presence of satellite peaks in CH$_3$CN in addition to D$_2$O suggests that the 1$^\text{st}$ reduction of ReBPMRubpyamide is partially accessible in both solvents studied without any potential difference applied. As before, it is reasonable to assume that the absence of such a satellite feature from the amide v(CO) region is due to the comparatively small amount of electronic coupling between the amide groups (situated on the bpy ligands) and the bipyrimidine bridge.
Figure 4.9.3.1: The FTIR (bottom) and ps-TRIR spectra (top) between 1 and 1000 ps (top) of the ReCO (left) and Rubpyamide (right) termini of ReBPMRubpyamide in D₂O solution.
Figure 4.9.3.2: ps-TRIR kinetics of ReBPMRubpyamide in D$_2$O solution for ReCO (transient at 2023 cm$^{-1}$, top left and bleach at 2028 cm$^{-1}$, bottom left) and Rubpyamide (transient at 1614 cm$^{-1}$, top right and bleach at 1635 cm$^{-1}$, bottom right) termini for $\Delta t = 0$-2000 ps.
Figure 4.9.3.3: The FTIR (bottom) and ns-TRIR spectra (top) between 1 and 250 ns (top) of ReBPMRubpyamide in D$_2$O solution.

Figure 4.9.3.4: ns-TRIR kinetics of bleach (1632 cm$^{-1}$, bottom) and transient (1614 cm$^{-1}$, top) bands of ReBPMRubpyamide in D$_2$O solution.
Picosecond and nanosecond-TRIR spectroscopy (with laser excitation at 400 and 355 nm respectively) of ReBPMRubpyamide were measured in CH₃CN solution, monitoring both regions in the IR where reporter groups are present.

**ps-TRIR Spectroscopy**

Figures 4.9.3.5 and 4.9.3.6 show the ps-TRIR spectra and kinetics obtained for the ReCO and Rubpyamide termini in CH₃CN.

**ReCO (“Donor” / Higher Energy) terminus**

The spectral profile obtained following 400 nm excitation is dominated by two bleaches at 2034 and 1936 cm⁻¹. Two transient bands are detected at lower energy to the bleaches, centred at 2023 and 1931 cm⁻¹. As in D₂O this is consistent with the ReCO terminus experiencing an increase in electron density as a result of a MLCT transition involving the Rubpyamide terminus (*vide infra*). There is considerable overlap between the bleaches and their associated transient bands. These bands narrow and blue-shift over the first 25 ps of the experiment as a result of vibrational cooling. A biexponential decay is observed with \( \tau_1 = 67 \) (±31) and \( \tau_2 = 970 \) (±150) ps (averaged over both bleach and transient features). At \( \Delta t = 1 \) ns only ca. 10% of the signal present at \( \Delta t = 1 \) ps remains. The similarity of these results to those obtained in D₂O is consistent with solvent-independence of the photophysics of ReBPMRubpyamide.

**Rubpyamide (“Acceptor” / Lower Energy) terminus**

The ps-TRIR spectra obtained are comprised of a bleach centred at 1639 cm⁻¹ and transient bands to lower and higher wavenumber centred at 1621 and 1650 cm⁻¹ respectively. The bipyrimidine feature, also detected in other Rubpyamide-containing complexes in this Chapter (*vide supra*) is observed at 1577 cm⁻¹ and its associated transient is at 1555 cm⁻¹. Examination of the kinetics of the amide bands shows evidence for a grow-in of absorbance over the first 75 ps on both bleach and transient features. As in D₂O, the timescale of this transient signal’s growth and its manifestation in both bleach and transient spectra indicates that vibrational cooling is
not responsible and there is a correlation with the fast decay process $\tau_1$ on the ReCO terminus (*vide supra*). After this growth has halted, the signal decays with a time constant (averaged over both bleach and transient features) of 917 ($\pm$88) ps. This is in very good agreement with the slow decay process $\tau_2$ on the ReCO terminus (*vide supra*). At $\Delta t = 1$ ns ca. 25% of the signal at $\Delta t = 1$ ps remains, indicating that a long-lived transient species persists into the nanosecond time domain.

**ns-TRIR Spectroscopy**

Figures 4.9.3.7 and 4.9.3.8 show the ns-TRIR spectra and kinetics obtained respectively for the Rubpyamide terminus in D$_2$O (no signal was detected in for the ReCO terminus). As with the ps-TRIR spectra (*vide supra*), the spectrum is dominated by a bleach centred at 1639 cm$^{-1}$ with transient bands at 1649 and 1621 cm$^{-1}$. The bipyrimidine mode bleach is detected at 1579 cm$^{-1}$ and its associated transient is centred at 1556 cm$^{-1}$. All the bands decay synchronously with a biexponential time constants (averaged over both bleach and transient features) of $\tau_1 = 1.4 \pm 0.1$ and $\tau_2 = 84.5 \pm 4.2$ ns. The $\tau_1$ ns-TRIR decay component shows relatively good agreement with $\tau_2$ in the Rubpyamide picosecond spectra (*vide supra*), considering that neither time domain is well calibrated for measuring lifetimes of a few nanoseconds.

**4.9.4. ReBPMRubpyamide: Closing Points**

The lowest energy excited state of ReBPMRubpyamide has been shown to be Ru $\rightarrow$ bpm $^3$MLCT, with transient bands from this state persisting into the nanosecond time domain. The spectral profiles observed are similar in both solvents, consistent with the expected lack of solvatochromism in the system. In contrast to ReBPMRuCN there is evidence in the picosecond TRIR kinetics for interconversion of Re $\rightarrow$ bpm into Ru $\rightarrow$ bpm $^3$MLCT excited states in both D$_2$O and CH$_3$CN. The satellite bands observed in the FTIR suggest that the 1$\text{st}$ redox couple near 0 V is accessible and this may provide a mechanism with which to switch on additional photoprocesses such as IVCT although there is no definitive evidence for this in this complex.
Figure 4.9.3.5: The FTIR (bottom) and ps-TRIR spectra (top) between 1 and 1000 ps (top) of the ReCO (left) and Rubpyamide (right) termini of ReBPMRubpyamide in CH$_3$CN solution.
Figure 4.9.3.6: ps-TRIR kinetics of ReBPMRpyamide in D$_2$O solution for ReCO (transient at 2022 cm$^{-1}$, top left and bleach at 2032 cm$^{-1}$, bottom left), ReCO (transient at 1630 cm$^{-1}$, top right and bleach at 1637 cm$^{-1}$, bottom right) termini for $\Delta t = 0$-2000 ps.
Figure 4.9.3.7: The FTIR (bottom) and ns-TRIR spectra (top) between 1 and 250 ns (top) of ReBPMRubpyamide in CH$_3$CN solution.

Figure 4.9.3.8: ns-TRIR single-pixel kinetics of bleach (1637 cm$^{-1}$, bottom) and transient (1651 cm$^{-1}$, top) bands of ReBPMRubpyamide in CH$_3$CN solution.
RuCNBPMPRubpyamide is a heterobimetallic complex with RuCN and Rubpyamide termini on either side of the bpm bridging ligand. As with the other heterobimetallic systems in this Chapter, each terminus has an M → bpm MLCT chromophore associated with it, and the Rubpyamide terminus possesses an additional Ru → bpyamide chromophore, which may complicate the photophysical dynamics of the system in a similar way to other Rubpyamide-containing complexes in this Chapter (§4.4, §4.7 and §4.10). This may become significant in the event of unselective excitation in the TRIR experiments. As in the RuCN-BL²-Rubpyamide complexes in Chapter 3, the RuCN terminus is strongly solvatochromic whilst the Rubpyamide terminus is almost solvent-independent. As a result the system’s photophysics is expected to show considerable differences in low and high A. N. solvents. In CH₃CN the RuCN → bpm transition is postulated to be the lowest energy excited state whilst in D₂O it is expected to be the Rubpyamide → bpm transition. This may provide an insight into the difference in electronic structure of the RuCN-BL²-Rubpyamide and RuCNBPMPRubpyamide systems, which are constructed from the same chromophores attached to two extreme cases of bridging ligands; strongly conjugated bpm and saturated BL¹ and BL².
Towards the end of the project it became apparent that the sample of RuCNBPMRubpyamide used in the experiments performed using the PIRATE apparatus at RAL (particularly affecting TRIR spectra collected in CH₃CN) contained a significant amount of NaCl impurity. Given other work in this thesis (§3.5.1) demonstrating the profound effect that alkali and alkaline earth metals have upon binding to the cyanide groups of the tetracyanoruthenate (II) unit it is conceivable that the bound Na⁺ ions have affected the Ru → bpm and Ru → bpyamide MLCT transition energies and the subsequent photophysics. We interpret the following anomalous results in terms of such metal binding.

4.10.1. UV/visible Absorption and Luminescence Spectroscopy

The UV/visible absorption spectra of RuCNBPMRubpyamide have been measured in D₂O and CH₃CN (see Appendix 4.12 for spectra). The results are presented in Table 4.10.1.1.

<table>
<thead>
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<th>Solvent</th>
<th>λ&lt;sub&gt;max&lt;/sub&gt;/nm</th>
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<tbody>
<tr>
<td>D₂O</td>
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<tr>
<td>CH₃CN</td>
<td>&lt; 300 364 402 457 494 750</td>
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</table>

Table 4.10.1.1: Table of peak positions and assignments for UV/visible absorption spectra of RuCNBPMRubpyamide in D₂O and CH₃CN solutions.

The emission spectrum of RuCNBPMRubpyamide was measured in CH₃CN solution at room temperature. No luminescence was detected.

4.10.2. Cyclic Voltammetry

Cyclic voltammetry of RuCNBPMRubpyamide was measured in CH₃CN (see Appendix 4.12 for voltammograms) and the results are presented in Table 4.10.2.1.
below. Features corresponding to three reversible reductions and three oxidation
couples (one reversible and two irreversible) are present. As with the other
heterobimetallics in this Chapter (§4.8, §4.9), the ease of reduction of the
bipyrimidine ligand in this case could be significant as the redox couple’s proximity
to 0 V indicates that this couple may be accessible in the absence of current, which
could have important implications for the electronic structure of the ground and
excited states of RuCNBPMRubpyamide in solution.

<table>
<thead>
<tr>
<th>$E_{re}/V$ (vs. Fe/Fe$^+$)</th>
<th>Reversibility</th>
<th>Tentative Assignment</th>
</tr>
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<tbody>
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<td>$E_p = +1.38$</td>
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<td>Re$^{II/III}$</td>
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</tr>
<tr>
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<td>Reversible</td>
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<td>Reversible</td>
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</table>

*Table 4.10.2.1: Table of potentials for redox couples and assignments from the
cyclic voltammogram of RuCNBPMRubpyamide in CH$_3$CN solution.*

4.10.3. FTIR, IR OTTLE and Time-Resolved Infrared Spectroscopy

**RuCNBPMRubpyamide** has been studied using ps and ns-TRIR spectroscopy in
D$_2$O and CH$_3$CN solutions. As a result of the negative solvatochromism of the RuCN
terminus we expect the excited state ordering to differ in the two solvents, as in the
RuCN-BL$^x$-Rubpyamide complexes in Chapter 3.

IR OTTLE of **RuCNBPMRubpyamide** in CH$_3$CN was unsuccessful as the 1$^{st}$
reduction couple was found to be chemically irreversible on the timescale of the
experiment.
Deuterium Oxide Solution

The FTIR spectrum of RuCNBPMRubpyamide in D$_2$O solution in the v(CN) and amide v(CO) region has been measured. In the cyanide region there is a peak profile characteristic of the tetracyanoruthenate (II) unit with 3 resolved maxima at 2099, 2064 and 2054 cm$^{-1}$. There is also a smaller ‘satellite’ peak visible at 2014 cm$^{-1}$, which may be composed of a number of unresolved bands. In the amide region there is a single peak present at 1613 cm$^{-1}$. These results are comparable with the FTIR spectra of the model complexes RuCNBPM and RubpyamideBPM (§4.2.3, §4.4.3). Satellite peaks at lower wavenumber to the cyanide bands has been observed and these are thought to be due to the partial accessibility of the complex’s 1$^{st}$ reduction in D$_2$O solution at 0 V. The peaks correspond to the FTIR spectrum of the partially reduced complex with the electron residing on the bipyrimidine ligand. No satellite bands have been observed in the amide v(CO) region of the spectrum for reasons discussed in §4.9.3.

Picosecond and nanosecond TRIR spectroscopy (with laser excitation at 400 and 355 nm respectively) were performed on RuCNBPMRubpyamide in D$_2$O solution, monitoring both regions in the IR where reporter groups are present. Each timescale will be discussed in turn and then the validity of the photophysical model will be evaluated.

ps-TRIR Spectroscopy

Figure 4.10.3.1 shows the ps-TRIR spectra obtained for the RuCN and Rubpyamide termini in D$_2$O. Figure 4.10.3.2 shows the ps-TRIR kinetics obtained for bleach and transient bands in D$_2$O.

RuCN (Higher Energy) Terminus

The spectral profile obtained following 400 nm excitation is dominated by a broad bleach with components centred at 2099, 2064 and 2054 cm$^{-1}$. A transient band is formed at higher energy to the bleaches, centred at 2102 cm$^{-1}$. This is consistent with
RuCN → bpm $^3$MLCT being the lowest energy excited state, as predicted by the photophysical model. A very weak transient band is also detected to lower wavenumber than the bleach at 1986 cm$^{-1}$. By invoking the dipole moment and IR oscillator strength arguments presented in §4.6.3, it is possible that this band may represent a very small population of molecules but does suggest that the excited state electronic charge resides on bpm and therefore that Ru → bpm MLCT is occurring. The bands decay monoexponentially with $\tau = 181$ (±24) ps (averaged over bleach and transient features). At $\Delta t = 1$ ns only ca. 5% of the signal at $\Delta t = 1$ ps remains. This indicates that quenching of the excited state is almost complete during the picosecond time domain.

Rubpyamide (Lower Energy) Terminus
The ps-TRIR spectra obtained are comprised of a parent bleach centred at 1613 cm$^{-1}$ and a transient band to lower wavenumber at 1601 cm$^{-1}$. This is consistent with an increase in electron density on the bpy ligands, expected during the formation of a Ru → ((4,4'-CONEt$_2$)$_2$bpy) $^3$MLCT excited state. This is in contradiction to the earlier statement in §4.1 that Ru → bpm $^3$MLCT transitions are always at lower energy than Ru → bpy $^3$MLCT transitions. Additionally a transient is observed at 1545 cm$^{-1}$, although no accompanying bleach is detected. This feature can be assigned to a ring mode of the bipyrimidine ligand, directly reporting on the increase in electron density on the bridging ligand.

Examination of the kinetics of the amide bands gives a biexponential relationship with decay constants $\tau_1 = 8$ (±3) and $\tau_2 = 190$ (±13) ps. There is good agreement of $\tau_2$ with the decay of the RuCN terminus. Interestingly the kinetics of the bipyrimidine transient band at 1545 cm$^{-1}$ exhibit a growth over the first 20 ps followed by decay with $\tau = 150$ (±15) ps. This growth gives good agreement with the fast decay component of the Rubpyamide terminus and strongly suggests that (as expected) the Ru → ((4,4'-CONEt$_2$)$_2$bpy) $^3$MLCT excited state is converting into the Ru → bpm $^3$MLCT excited state.
If this was the case, then we would expect to observe transient features at higher wavenumber to the bleach in the amide spectrum (as in ReBPMRubpyamide, §4.9.3). However these bands are not observed. This anomalous result may be a consequence of NaCl impurity in the sample perturbing the electronic structure and photophysical behaviour of the complex via Na$^+$ binding to the cyanide ligands.

**Acetonitrile Solution**

**ps-TRIR Spectroscopy**

Figure 4.10.3.3 shows the ps-TRIR spectra obtained for the ReCO and Rubpyamide termini in CH$_3$CN. Figure 4.10.3.4 shows the ps-TRIR kinetics obtained for bleach and transient bands in CH$_3$CN.

**RuCN (Lower Energy) Terminus**

The spectral profile obtained following 400 nm excitation is dominated a bleach with components centred at 2099, 2075 and 2065 cm$^{-1}$. A broad transient band is detected at lower energy to the bleach, centred at 2033 cm$^{-1}$. This is consistent with the RuCN terminus experiencing an increase in electron density due to $^{3}$MLCT excited state formation on the other terminus. No transient band was detected to higher wavenumber, although the data obtained suffer from noise due to instability of the probe laser at RAL and weak bands may be indistinguishable from baseline fluctuations. The reader is reminded of the *caveat* with this system (particularly in CH$_3$CN) as a result of sample impurity (§4.10.1). A biexponential decay is observed with $\tau_1 = 16$ (±8) and $\tau_2 = 96$ (±13) ps (averaged over both bleach and transient features). At $\Delta t = 1$ ns less than 5% of the signal at $\Delta t = 1$ ps remains, reiterating the very short lifetime of the excited state formed.
Figure 4.10.3.1: The FTIR (bottom) and ps-TRIR spectra (top) between 1 and 1000 ps (top) of the RuCN (left) and Rubpyamide (right) termini of ReBPMRubpyamide in D$_2$O solution.
Figure 4.10.3.2: ps-TRIR kinetics of RuCNBPMP rubpyamide in D₂O solution for RuCN (transient at 2001 cm⁻¹, top left and bleach at 2065 cm⁻¹, bottom left) and Rubpyamide (transient at 1599 cm⁻¹, top right and bleach at 1618 cm⁻¹, bottom right) termini for Δt = 0-1000 ps.
Rubpyamide (Higher Energy) Terminus

The ps-TRIR spectra obtained are comprised of a bleach centred at 1637 cm\(^{-1}\) and a transient band to lower wavenumber centred at 1623 cm\(^{-1}\), implying that a Ru→((4,4'-CONEt\(_2\))\(_2\)bpy)\(^3\)MLCT excited state is formed, again in contrast to the model's prediction that Ru→bpm\(^3\)MLCT would be the lowest excited state formed. No bands arising from bpm modes are observed in this spectrum. The bleach and transient bands decay rapidly with \(\tau = 13 (\pm 1)\) ps, synchronous with \(\tau_1\) of the RuCN terminus.

These results suggest that the Rubpyamide terminus is lower in energy in D\(_2\)O and the RuCN terminus is lower in energy in CH\(_3\)CN, as predicted by the model. But our photophysical model is unable to explain the evidence from kinetic analysis of the ps-TRIR spectra in CH\(_3\)CN (particularly of the bpm vibrational mode) the apparent excited state interconversion of Ru→bpyamide\(^3\)MLCT into Ru→bpm\(^3\)MLCT without any changes accompanying the profile of Rubpyamide terminus. The insensitivity of the amide bands to the change of location of the excited state is perplexing as in all other complexes studied in this Thesis the amide band has proved to be a sensitive probe of excited state electronic structure. The only explanation offered is that impurities in the samples used are responsible for the erroneous results shown here.

4.10.4. RuCNBPMPrubpyamide: Closing Points

As predicted the strong solvatochromism of RuCNBPMPrubpyamide has led to a reversal in excited state ordering in low and high A. N. solvents. In D\(_2\)O the Rubpyamide terminus is the lower in energy whilst the RuCN terminus is lowest in CH\(_3\)CN. The similar excited state energies of the two termini leads to very short-lived excited states; higher lying but thermally accessible excited states may be responsible for additional rapid deactivation pathways. As with ReBPMPRuCN there is no direct kinetic evidence for interconversion of excited states on the picosecond timescale but these processes may be occurring on the femtosecond timescale.\(^{21}\)
Figure 4.10.3.3: The FTIR (bottom) and ps-TRIR spectra (top) between 1 and 1000 ps (top) of the RuCN (left) and Rubpyamide (right) termini of ReBPMRubpyamide in CH₃CN solution.
Figure 4.10.3.4: ps-TRIR kinetics of RuCNBPMPrubpyamide in CH$_3$CN solution for RuCN (transient at 2033 cm$^{-1}$, top left and bleach at 2073 cm$^{-1}$, bottom left) and Rubpyamide (transient at 1624 cm$^{-1}$, top right and bleach at 1637 cm$^{-1}$, bottom right) termini for $\Delta t = 0$-1000 ps.
Many of the TRIR spectra and kinetic traces measured for \textbf{RuCNBPMRubpyamide} give inconsistent or anomalous results compared to other complexes in this Chapter. The reader is reminded of the sample contamination that was discovered after the experiments were conducted which may be responsible for the unexplained phenomena observed.

As with the other heterobimetallic systems in this Chapter, another contributing factor to the complexity of the ps-TRIR kinetics may arise from the presence of additional \textsuperscript{3}MLCT states involved in the excited state relaxation process. These states are likely to possess slightly higher energies than the lowest-lying one due to the presence of two $\pi^*$ LUMOs in the bpm ligand with similar energies (Figure 4.1.1.2). Additionally the high sample concentrations required experimentally to obtain good signal-to-noise ratios with the PIRATE apparatus may lead to self-quenching and/or aggregation (see Chapter 6 for a discussion of this effect).

4.11. Chapter Discussion and Closing Points

In this Chapter, previous studies of d\textsuperscript{6} transition metal complexes have been presented and the limitations of optical spectroscopic techniques such as UV/vis and luminescence in determining the excited state electronic structure of bpm complexes have been discussed. A series of monometallic, homobimetallic and heterobimetallic complexes based on the building blocks [ReCl(CO)\textsubscript{3}], [Ru(CN)\textsubscript{4}]\textsuperscript{2-} and [Ru(4,4-(CONEt\textsubscript{2})\textsubscript{2}bpy)\textsubscript{2}]\textsuperscript{2+} and the bidentate bridging ligand 2,2'-bipyrimidine (bpm) were prepared and their spectroscopic and electrochemical properties studied.

The monometallic complexes \textbf{ReBPM, RuCNBPM} and \textbf{RubpyamideBPM} each possess \textsuperscript{3}MLCT lowest-lying excited states and were studied using TRIR on the picosecond and nanosecond timescales, a substantial improvement on the amount of information previously available using UV/vis absorption and emission spectroscopy. The excited state energies of these complexes are found to be lower than corresponding complexes of 2,2'-bipyridine, in agreement with previous reports and
this may be a contributing factor to the inherently weak luminescence of bpm-containing complexes. In the case of RubpyamideBPM the presence of two polypyridyl chromophores gives rise to more complicated photophysical dynamics, with excited state interconversion of Ru → bpy 3MLCT into Ru → bpm 3MLCT observed using the amide ν(CO) bands and vibrational modes of the bpm ligand. The results obtained were found to be similar to those in the literature for identical or related complexes.

The homobimetallic complexes (Re)2BPM, (RuCN)2BPM and (Rubpyamide)2BPM display more complex photophysics than their monometallic precursors. For the pair (Re)2BPM and (RuCN)2BPM whose IR reporters are ligands directly connected to the metal centre, the ps-TRIR spectra collected indicate that there is electronic asymmetry in the excited state; one metal centre is undergoing MLCT excitation to the bridging bpm and the other is sensing it, giving rise to two sets of excited state bands. These are located either side of the ground state bleach, such that one set of bands are shifted to higher wavenumber and the other set are shifted to lower wavenumber relative to the parent bands. These results indicate that the excited states probed are localised on the IR timescale (§1.3), in keeping with similar studies in the literature (§4.1.3).10 (Rubpyamide)2BPM's IR reporter groups are attached to pendant bpy ligands and as such return less detailed spectroscopic information on the electron density at the metal centre, however insights on excited stated interconversion have been gleaned from the kinetic data collected and investigation of the lineshapes of the spectra (vide infra). The results obtained for (Re)2BPM and (Rubpyamide)2BPM were found to be similar to those in the literature for identical or related complexes.

It is interesting to compare the lifetimes of ReBPM and (Re)2BPM vs. RuCNBPM and (RuCN)2BPM in CH3CN. Although in both cases the dinuclear complexes have a shorter lifetime than the corresponding mononuclear complexes, the difference is significantly less between RuCNBPM and (RuCN)2BPM (τMono = 3 × τBimetallic) than between ReBPM and (Re)2BPM (τMono = 26 × τBimetallic). Secondly, although the

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lifetimes of both RuCNBPM and (RuCN)2BPM in D2O are much longer than in CH3CN. the same ratio of lifetimes is observed (τMono ≈ 3 × τ Bimetallic). This can be ascribed to the fact that addition of a second [ReCl(CO)3] centre (on changing from ReBPM to (Re)2BPM) reduces the lowest 1MLCT absorption by 5400 cm⁻¹, whereas addition of a second [Ru(CN)4]²⁻ centre (on changing from RuCNBPM to (RuCN)2BPM) has a smaller effect on the lowest MLCT state, reducing the lowest 1MLCT absorption by only 3300 cm⁻¹ in water. Consequently there is a bigger difference between the LUMO energies of ReBPM and (Re)2BPM than between RuCNBPM and (RuCN)2BPM, and the energy-gap law will result in a bigger drop in the excited state lifetimes between ReBPM and (Re)2BPM. It is likely that the weaker stabilising effect on the bpm LUMO of an additional [Ru(CN)4]²⁻ unit compared to an additional [ReCl(CO)3] is because the former is a poorer π-acceptor on account of its doubly negative charge.

Lineshape analysis of the ps-TRIR spectra of RubpyamideBPM and (Rubpyamide)2BPM determined that there is an obscured transient feature at lower energy to the bleach for RubpyamideBPM but not for (Rubpyamide)2BPM. This suggests that the Ru → bpy into Ru → bpm 3MLCT excited state conversion process occurs on the picosecond timescale in the monometallic complex but on the femtosecond timescale in the homobimetallic complex.

The heterobimetallic complexes ReBPMRuCN, ReBPMRubpyamide and RuCNBPMRubpyamide each possess two different chromophores and IR reporter groups, generally resulting in more complicated photophysical dynamics than the homobimetallic complexes. In the case of ReBPMRuCN there is substantial overlap between the IR bands of the RuCN and ReCO termini, resulting in a very convoluted spectrum which is difficult to meaningfully analyse. The spectra obtained for ReBPMRubpyamide indicate that excited state conversion from Re → bpm 3MLCT into Ru → bpm 3MLCT is observed on the picosecond timescale. RuCNBPMRubpyamide has an interesting relationship with the RuCN-BLx-Rubpyamide systems in Chapter 4. The replacement of the saturated bridging ligand
BL\textsuperscript{1}/BL\textsuperscript{2} for the conjugated bpm ought to have provided an interesting comparison, but the spectra collected for RuCNBPM\textsubscript{2}Rubpyamide are somewhat anomalous and provide limited insight due to a problem with NaCl impurity. However this is ostensibly a testament to the strong binding of alkali and alkali earth metals to cyanide ligands, an effect which is deliberately exploited in Chapter 3!

In both heterobimetallic complexes containing a ReCO terminus the initial laser excitation is unselective, populating \(^3\text{MLCT}\) excited states on either terminus. Internal relaxation processes then facilitate Re \(\rightarrow\) bpm to Ru \(\rightarrow\) bpm \(^3\text{MLCT}\) excited state interconversion, with evidence for this from the grow-in of absorbance of the Rubpyamide terminus of ReBPM\textsubscript{2}Rubpyamide in D\textsubscript{2}O and CH\textsubscript{3}CN. The polyexponential time dependence of heterobimetallic ps-TRIR has been rationalised in terms of the presence of (at least) two independent \(^3\text{MLCT}\) states involved in the excited state relaxation process. These states are likely to possess similar energies due to the presence of two \(\pi^*\) LUMOs in the bpm ligand with similar energies (Figure 4.1.1.2).

In the complex RuCNBPM\textsubscript{2}Rubpyamide which presented some anomalous results, ps-TRIR suggests a similar scenario occurs with unselective excitation and subsequent excited state relaxation to the lowest-energy \(^3\text{MLCT}\). The unusually short lifetimes of the excited states formed may be in part due to a number of factors, discussed in §4.10.4. These include the projected proximity of the energies of the lowest \(^3\text{MLCT}\) excited states located on the RuCN and Rubpyamide termini leading to additional thermally-accessible excited state deactivation pathways, multiple close-lying \(^3\text{MLCT}\) excited states on each terminus (due to the electronic structure of bpm, \textit{vide supra}) and the contamination of the sample with NaCl impurity. The latter is likely to result in the binding of Na\textsuperscript{+} ions to the cyanide ligands – an effect which demonstrably alters the electronic structure of the RuCN terminus to the extent that excited state ordering may be affected (§3.5.1).
Although the possibility of IVCT transitions in these complexes has been mentioned at several points throughout the Chapter on account of the extent of electronic communication between the two metal centres in such bimetallic systems, no evidence for such a process has been identified. The ground state structures of these complexes are unlikely to exhibit IVCT – a redox event may be required to generate a mixed-valence species (e.g. Ru\textsuperscript{II}-bpm-Ru\textsuperscript{III}) which may then undergo IVCT processes. It was postulated that the presence of a redox couple in the cyclic voltammograms of some bimetallic complexes near 0 V (see §4.12 for the voltammograms obtained) may have provided the necessary ‘switching’ event to activate IVCT but no evidence has been observed for this in terms of low energy UV/Vis absorption bands or nIR luminescence features.

Finally, returning to the questions posed at the start of this Chapter:

- Rhenium (I) and ruthenium (II) complexes of bpm exhibit weaker luminescence and shorter excited state lifetimes than the corresponding bpy complexes because the bpm ligand possesses multiple low energy LUMOs which in accordance with the energy-gap law the smaller HOMO-LUMO energy difference leads to short-lived excited states and more efficient non-radiative decay.

- All of the bimetallic complexes have been shown to possess excited states which are localised on the IR timescale. In many cases there are multiple excited state species present in solution and efficient interconversion to form the lowest energy \(^3\)MLCT excited state occurs on the femtosecond or picosecond timescales. The observation of many of these ‘dark’ (non-luminescent) states would not be possible using emission spectroscopy.

- The consequences of unselective excitation of MLCT chromophores have been observed with some convoluted spectra (e.g. ReBPMRuCN) from which it has been difficult to extract information from. It is conceivable that higher
resolution TRIR spectroscopy would be useful to understand such systems further.

- No evidence for IVCT processes has been observed using UV/vis absorption, luminescence or TRIR spectroscopy or electrochemistry.
4.12. Appendix of Spectra and Voltammograms
Measurements on ReBPM in CH₂CN: UV/Vis spectrum (left); luminescence spectrum (λex = 375 nm, centre); cyclic voltammograms (right).

UV/Vis spectrum of RuCNBP in D₂O (black) and CH₃CN (red).
RubpyamBPM

Measurements on RubpyamBPM: UV/Vis spectrum (left) in D₂O (black) and CH₃CN (red); luminescence spectrum in CH₃CN (λₑₓ = 461 nm, centre); cyclic voltammogram n CH₃CN (right).

(Re)₂Bpm

UV/Vis spectrum of (Re)₂BPM in CH₃CN.
**Measurements on (RuCN)$_2$BPM:** UV/Vis spectrum (left) in D$_2$O (black) and CH$_3$CN (red); cyclic voltammogram in CH$_3$CN (middle and right).

**Measurements on (Rubpyam)$_2$BPM:** UV/Vis spectrum (left) in D$_2$O (black) and CH$_3$CN (red); luminescence spectrum in CH$_3$CN ($\lambda_{ex} = 425$ nm, black and $\lambda_{ex} = 550$ nm, red, centre); cyclic voltammogram in CH$_3$CN (right).
ReBPMRuCN

Measurements on ReBPMRuCN: UV/Vis spectrum (left) in D₂O (black) and CH₃CN (red); cyclic voltammogram in CH₃CN (right).

ReBPMRubpyamid

Measurements on ReBPMRubpyamide: UV/Vis spectrum (left) in D₂O (black) and CH₃CN (red); luminescence spectrum in CH₃CN (λₑₓ = 425 nm, black and λₛₑₓ = 550 nm, red, centre); cyclic voltammogram in CH₃CN (right).
Measurements on RuCNBPMPRubpyamide: UV/Vis spectrum (left) in D$_2$O (black) and CH$_3$CN (red); cyclic voltammogram in CH$_3$CN (right).
4.13. References

Chapter 5:

Closing Points and Future Work
Chapter 5: Closing Points and Future Work

5.1. Introduction

In this Thesis, "Solvent Based Switching of the Photophysical Properties of Polynuclear Complexes" two different types of photophysically active supramolecular system have been designed and studied using a variety of spectroscopic, electrochemical and theoretical techniques, in low and high Acceptor Number solvents (CH$_3$CN and D$_2$O respectively). The transition metal moieties used as modular building blocks, [Ru(CN)$_4$]$^{2-}$, [Ru((4,4'-CONEt$_2$)$_2$bpym)$_2$]$^{2+}$ and [ReCl(CO)$_3$] were combined with the bridging ligands BL$^1$ and BL$^2$ (Chapter 3) and bpm (Chapter 4) to construct monometallic$^1$, homobimetallic$^2$ and heterobimetallic$^3$ complexes containing zero, one or two strongly solvatochromic [Ru(CN)$_4$]$^{2-}$ moieties. This has enabled the study of 'solvent-switchable' systems (e.g. RuCN-BL$^1$-Rubpyamide, §3.5) to be carried out alongside bimetallic complexes which exhibit strong negative solvatochromism but no solvent dependence in the location of their lowest $^3$MLCT excited state (e.g. (RuCN)$_2$BPM, §4.6) and also complexes which exhibit strong solvatochromism but the same excited state ordering regardless of solvent environment (e.g. ReBPMRuCN, §4.8).

A study of tetracyanoruthenate (II) complexes with a series of polypyridyl ligands was presented in Chapter 2, providing a valuable dataset for the scientific community with UV/vis, luminescence and TRIR spectroscopies, crystallography$^*$ and electrochemistry. This study has demonstrated the utility of the v(CN) bands of the [Ru(CN)$_4$]$^{2-}$ unit to act as a sensitive reporter of ground and excited state electron density.

Furthermore an additional mechanism for modulating the $^3$MLCT excited state energy of the [Ru(CN)$_4$]$^{2-}$ unit was also briefly demonstrated by way of a luminescence titration of RuCN-BL$^1$-Rubpyamide with Ba$^{2+}$ ions (§3.5.1). This study has recently

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* Performed at the University of Sheffield by Mr. T. L. Easun and co-workers and not shown in this Thesis.
been expanded upon by Prof. M. D. Ward and co-workers through spectroscopic monitoring of titrations with a variety of metal ions and tetracyanoruthenate (II) complexes and is the subject of a recent publication. This article (in conjunction with the luminescence titration detailed in §3.5.1) demonstrates the ability of the cyanide ligands of the [Ru(CN)₄]²⁻ unit to strongly and reversibly bind metal ions which modulate the ³MLCT excited state energies of the [Ru(CN)₄(NN)]²⁻ moiety.

The ps and ns-TRIR spectra shown throughout this Thesis which have been central to probing the photophysical processes occurring in Chapters 2, 3 and 4 were recorded using the PIRATE spectrometer at the Rutherford Appleton Laboratory. This meant that the amount of time which could be devoted to collecting spectra and designing experiments based on the results obtained was largely determined by our success (or otherwise) in being awarded experimental time at the Central Laser Facility and various experimental issues such as the performance of the pump and probe lasers, detectors and data acquisition software. Our proposals to study our solvent-switchable bimetallic systems (Chapters 3 and 4) using ultrafast transient UV/vis and Raman spectroscopies were unsuccessful, restricting us to focus on ultrafast TRIR spectroscopy as our primary time-resolved technique.

With the imminent completion of a custom-built picosecond time-resolved IR, UV/vis, Raman and fluorescence spectroscopic suite in the School of Chemistry, University of Nottingham more detailed and complete investigations will be possible. Such investigations could include more advanced experiments such as t2D-IR hole-burning and labelling experiments (§1.4.1.3). With this in mind one hopes that the next generation of Nottingham spectroscopists will be able to use these state-of-the-art facilities to design and execute more complete ultrafast spectroscopic studies of transition metal excited states.
5.2. Closing Points and Future Work

In Chapter 2 "Tetracyanoruthenate (II) Complexes With α-Diimine Ligands" a family of $[\text{Ru(CN)}_4(\text{NN})]^2-$ complexes was presented and studied. The dataset presented (and also those in Chapters 3 and 4) would benefit from study of the photophysical properties (and hence the effect of SSDA interactions) in a greater number of solvent environments. A good candidate for such an environment might be CH$_3$OH. This would be a suitable choice as its A. N. (41.3) is approximately midway between that of D$_2$O (54.8) and CH$_3$CN (19.3). Unfortunately several strong solvent peaks in the ν(CN) region of the infrared spectrum of CH$_3$OH precluded its use as a solvent in our TRIR investigations. It is interesting to note that Indelli et al. used CH$_3$OH to demonstrate an interesting case of equilibrium between the excited states of two chromophores in a supramolecular system based on $[\text{Ru(CN)}_4(\text{bpy})]^2-$ and pyrene (§3.1) using luminescence spectroscopy. As an aside, some ps-TRIR experiments were attempted on 'Indelli-type' complexes but were unsuccessful due to insufficient solubility for TRIR experiments. An alternative approach could be to use an equimolar solvent mixture of D$_2$O and CH$_3$CN to give an Acceptor Number intermediate between those of the pure solvents (37.1).

The close proximity of the four ν(CN) ground state bands of $[\text{Ru(CN)}_4(\text{NN})]^2-$ complexes (typically within 50 cm$^{-1}$ of each other) and the heavy overlap between bleach and transient bands, limited the amount of information which could be extracted from the cyanide region of TRIR spectra of systems incorporating the $[\text{Ru(CN)}_4]^2-$ unit. This was further compounded by the low spectral resolution of the PIRATE spectrometer (ca. 8 - 9 cm$^{-1}$) when configured for a wide enough spectral range to collect the entire ν(CN) spectrum over the 64 detector pixels. An attempt was made to simulate a higher resolution TRIR spectrum by splicing together two low resolution spectra (§2.3). However no major insights were gained, and it is unclear if this is because of feature obscuration as a result of spectral overlap or imprecision in the combination of the constituent spectra. In either case the acquisition of higher resolution TRIR (2 - 4 cm$^{-1}$) spectra in the ν(CN) region would no doubt further
spectroscopic understanding of the relative ordering of excited state $v$(CN) bands in systems containing the $\text{[Ru(CN)₄]}^{2-}$ unit.

An interesting point was also raised following the discussion of the recent publication by Kettle et al. regarding the possible significance of quadrupolar interactions in affecting the analysis of vibrational spectra of metal cyanide complexes, in addition to the well-known effect of dipole moment changes on vibrational spectra.$^7$ At this stage it is not clear how such second-order interactions could best be studied and quantified, but this effect may be associated with the wide variations in IR band intensity observed for different metal centres in the ps-TRIR spectra of $(\text{RuCN})₂\text{BPM}$ ($\S$4.6.3). Perhaps the combination of high-resolution FTIR, ps-TRIR and Levenberg-Marquardt lineshape fitting routines, similar to those used in a previous investigation by the author to detect dipole-octupole interactions in thin solid films might offer some insight.$^8$

In Chapter 3 "Solvent Based Switching of the Gradient and Direction of Förster Photoinduced Energy Transfer in Polynuclear Complexes", two termini based on ruthenium (II) and rhenium (I) centres were linked via a saturated CH₂CH₂ bridge connecting bipyridyls ligated on either metal centre. The complexity of the kinetics obtained from TCSPC and TRIR spectroscopy suggested that there were multiple photophysical processes occurring in solution, including through-space PEnT.$^9$ A more rigorous data fitting procedure might reveal more precisely (or at least give a 'ballpark' estimate of) the number of discrete components present in the multiexponential traces obtained.

In the systems designed to exhibit solvent-switchable photophysics (RuCN-BLₓ-Rubpyamide) the ps-TRIR kinetic information on the $\Delta t = 0-50$ ps timescale provided direct evidence supporting the photophysical model proposed by Simpson et al. ($\S$3.1).$^{10}$ A major triumph of this work is the direct observation of PEnT using vibrational spectroscopy in real time in the picosecond timescale.
However the kinetic data from TRIR spectroscopy were not totally consistent over both termini in many cases; in particular some (single-pixel) traces from the Rubpyamide terminus were ambiguous and exhibited neither a grow-in or a decay. It is conceivable that longer data acquisition times and improved band fitting could lead to better quality TRIR spectra and kinetics and it would be interesting for these systems to be reinvestigated using the new apparatus in Nottingham. In combination with a ps-TRIR study of Simpson's complex RuCN-BL-Rubpy a more detailed picture of the similarities and differences between the two systems (discussed in §3.8) might emerge.

There is some evidence from ps-TRIR spectroscopy for a 2-step PEnT process in RuCN-BL$^1$-Rubpyamide in D$_2$O (§3.5.3). However, more evidence is required to gain a better understanding of this effect, in particular the reasons for its absence in RuCN-BL$^2$-Rubpyamide in D$_2$O, which is thought to possess termini with almost identical energetics. The key differences between these two complexes are ostensibly conformational, however only a rudimentary conformational analysis was conducted in Chapter 3. A possible avenue for future extensions to this work could involve experimental studies of the luminescence anisotropy in conjunction with a thorough computational study of the conformations likely present in solution. Sophisticated algorithms used for protein folding, such as Stanford's recent distributed computation project 'Folding@Home', could provide a weighted distribution of the ensemble of supramolecular conformers likely to be present for the different systems in various solvents, taking into account the electrostatic properties of the two termini.

More generally an in-depth ground and excited state Density Functional Theory (DFT) study of [Ru(CN)$_4$(bpy)]$^2^-$, which to date has only been studied in brief, and polynuclear complexes based on it would add considerable gravitas to the analysis of the spectroscopic results presented in this Thesis. Such a study was attempted using the GAUSSIAN03 computational suite. However, the complexity of the polynuclear systems studied and the unusual nature of the SSDA interactions (which are responsible for these complexes' interesting photophysical behaviour) could not be
satisfactorily modelled using the simple ‘out-of-the-box’ algorithms and the author’s limited knowledge of DFT and computational chemistry. The University of Nottingham has recently installed the GAUSSIAN software on its supercomputer and an interesting avenue for future work in this area would be to complement the experimental ground and excited state spectroscopic results with a comprehensive DFT study.

It is interesting to consider that the perceived reduction in conformational degrees of freedom from BL\textsuperscript{1} to BL\textsuperscript{2} (i.e. from an infinite ensemble to two fairly rigid conformers) did not in fact result in a reduction in the number of exponents in the kinetics obtained by luminescence and TRIR spectroscopy. The reasons for this are unclear but the possibility of additional mono and bimetallic species' presence as impurities produced as a result of the synthesis of the BL\textsuperscript{2} ligand was raised in §3.4. The difference in photophysical behaviour between the solvatochromic RuCN-BL\textsuperscript{1}-Rubpyamide and ReCO-BL\textsuperscript{1}-Rubpyamide which is a solvent-independent system, is also not fully understood. The timescale of PEnT in the solvent-switchable systems containing a RuCN terminus (§3.5.3, 3.6.3) was on the Δt = 0-10 ps timescale, whilst in the system containing a ReCO terminus (§3.7.3) this process took place on the Δt = 0-5 ns timescale. Energy differences between the $^3$MLCT excited states on the ReCO and Rubpyamide termini are thought to be important factors in addition to the different electrostatic properties of the two termini. However the $^3$MLCT energies of the two termini in each of the systems was calculated to be relatively similar based on luminescence energies of the model complexes RuCNdmb, Rubpyamide-BL\textsuperscript{1} and [ReCl(CO)(bpy)] (§1.4.1.2, 3.2.1, 3.3.1, 3.4.1).

The rates of interchromophoric PEnT in Chapter 3 were not measured sufficiently accurately to calculate meaningful Förster parameters, in particular $\kappa^2$, the orientational term. There is scope for the collection of more complete datasets to explore and calculate these parameters, possibly also using ultrafast time-resolved UV/vis. Such datasets might include more ps-TRIR time delays, higher resolution spectra and more advanced fitting routines. This might facilitate the calculation of the
rate of through-space PEnT and the ‘effective’ interchromophoric distance, often used as a ‘molecular yardstick’ to gauge distances between chromophores in non-rigid systems.\textsuperscript{14}

In Chapter 4 “Probing the Excited States of Ruthenium (II) and Rhenium (I) Complexes of 2,2’-Bipyrimidine using TRIR Spectroscopy”, two metal centres were linked via the conjugated bpm ligand, with the two metal centres held in close proximity and with electronic interaction between them in the ground and excited states. One of the key properties of the systems studied was the timescale of excited state delocalisation, and TRIR spectroscopy is perhaps not the most suitable technique to probe this as the characteristic timescale of IR spectroscopy is on the femtosecond timescale. The use of additional spectroscopic techniques such as femtosecond and picosecond time-resolved UV/vis spectroscopy might facilitate direct investigation of the timescale of excited state localisation/delocalisation, and would also enable observation of the interconversion of Ru → bpy into Ru → bpm \textsuperscript{3}MLCT in bimetallic complexes in the cases when this process was apparently complete within 1 ps. Time-resolved UV/vis might also allow further study of IVCT processes in the heterobimetallic complexes \textit{ReBPMRuCN}, \textit{ReBPMRubpyamide} and \textit{RuCNBPMRuCN}. The TRIR spectra obtained for \textit{ReBPMRuCN} (§4.8.3) were somewhat convoluted due to the presence of approximately twenty $\nu$(CN) and $\nu$(CO) bands between 2100 and 1850 cm\textsuperscript{-1}, arising from the two excited state species present in solution following excitation. A proposal for experimental time at RAL to study these systems using time-resolved UV/vis and Raman spectroscopies was unsuccessful but these experiments could be carried out using the new facility in Nottingham.

A possible extension to the study of bipyrimidine-linked systems in Chapter 4 is to synthesise and study polynuclear complexes linked by cyanide ligands using ultrafast TRIR spectroscopy. There are a number of reports describing such systems in the
literature (§1.4.2.2) and would provide an interesting comparison with bipyrimidine-linked systems.

The apparently anomalous results obtained for RuCNBPMR ubpyamide, which was explained in terms of Na\(^+\) binding to the cyanide ligands, requires further study to verify the cause. In an analogous manner to the Ba\(^{2+}\) titrations in §3.5.1, a ps-TRIR titration experiment with Na\(^+\) ions and a macrocycle with appropriate cavity size to sequester the ions e.g. a cryptahemispherand\(^{15}\), might be helpful in confirming the nature of the impurity/anomaly more precisely. Atomic absorption spectrometry may also be of use in this case to directly measure the sodium ion concentration. Unfortunately a luminescence titration would not be suitable due to the absence of significant emission from the system (§4.11.1).

Further studies are required to investigate the significance of the redox processes apparently occurring at 0 V observed in cyclic voltammograms of homobimetallic and heterobimetallic bpm complexes. A combination of UV/vis OTE and EPR may prove useful in conclusively verifying whether these processes are occurring, and if they have any effect on the ground and excited state properties of the complexes studied in Chapter 4.

The particularly strong solvatochromism of (RuCN)\(_2\)BPM, which results from its eight cyanide ligands (§4.6), provides scope for studies designed to exploit this. Potential applications include incorporating the complex into a moisture, solvent or metal ion sensor (directly exploiting the colour change in the presence of solvent molecules or ions) or as a photosensitiser and charge injector into nanocrystalline TiO\(_2\), taking advantage of its broad \(^1\)MLCT absorption profile (covering most of the visible spectrum) and strongly variable \(^3\)MLCT excited state energies.

An interesting consideration is that the intercomponent photophysical processes occurring in bimetallic complexes of bipyrimidine may be described with equal validity as excited state electron or energy transfer. As the process occurring is
identical in either case, it is a semantic argument as to which describes the process most accurately. This has been discussed in a previous study.\textsuperscript{16}

Taking a wider view of the project, this Thesis, and possible future directions, a long-term goal of the scientific community is to incorporate systems such as the ones discussed in Chapters 3 and 4 into proto-devices like those envisioned by proponents of 'supramolecular machines' such as Lehn, Balzani and Stoddart.\textsuperscript{17} Considering the systems described in this Thesis as potential device components one can foresee the construction of light-driven supramolecular machines\textsuperscript{18} whereby intercomponent energy and electron transfer processes may occur in the same system. The dominant excited state deactivation pathway would then depend on some external stimulus e.g. solvent or atmospheric environment or the presence of metal ions which might reversibly bind to the cyanide ligands of the \([\text{Ru(CN)}_4(\text{NN})]^2-\) unit, modulating the energetics of the lowest energy \(^3\text{MLCT}\) excited state.

Systems incorporating the \([\text{Ru(CN)}_4(\text{bpy})]^2-\) unit and a cobaltocene moiety (\textbf{RuCN-Cob}, see Figure 5.2.1) were designed to study solvent-dependent PET and were analysed using luminescence and TRIR spectroscopy. However, no meaningful results were obtained. A more detailed study of such systems would be of considerable interest for the molecular electronics community, as the facile control of direction and thermodynamic gradient of PET would be an important step towards the construction of supramolecular device components acting as electronic switches (\textit{i.e.} molecular transistors).\textsuperscript{19}

![Figure 5.2.1: The RuCNbpy-Cob complexes studied unsuccessfully.](image-url)

5-9
The construction of higher nuclearity transition metal cyanide complexes, which was explored briefly during the course of this project using the ligand HAT (not shown in this Thesis) would offer additional benefits over the systems described here. These include a higher degree of connectivity to a single node and a potentially greater degree of modulation of their photophysical properties via solvatochromism and metallochromism phenomena, since these depend on the number of cyanide ligands present.

We recently studied the series of complexes \([\text{Ru}(\text{CN})_4(\text{HAT})]^{2-}\), \({\{\text{Ru}(\text{CN})_4\mu_2(\text{HAT})\}}^{4-}\) and \({\{\text{Ru}(\text{CN})_4\mu_3(\text{HAT})\}}^{6-}\) in which one, two or three \([\text{Ru}(\text{CN})_4]^{2-}\) units are connected to the bidentate sites of the tridentate ligand hexaaza-triphenylene (HAT).\(^{20,21}\) These systems would be expected to show a near-linear relationship with degree of solvatochromism and number of cyanide ligands (§2.1.3). Therefore \({\{\text{Ru}(\text{CN})_4\mu_3(\text{HAT})\}}^{6-}\) would be expected to exhibit the most marked degree of negative solvatochromism yet observed in complexes of this type owing to its twelve cyanide ligands. This would suggest that this complex would show even greater promise as a solvent/moisture/metal ion sensor that the bimetallic complex \((\text{RuCN})_2\text{BPM}\).

Another potentially interesting avenue of research identified is the broadening of the scope of the metal centres used in this Thesis. Extending the study to metals such as osmium (II), rhodium (I/III) and iridium (I/III) and platinum (II) would not only increase the potential synthetic functionalisation, design flexibility and utility of such systems, but would also ostensibly open up the possibility of device components exploiting a ‘two-tier’ solvatochromic effect. Two different metal cyanide supramolecular building blocks in the same system could exhibit varying degrees of negative solvatochromism such that there may be more than one Acceptor Number switching point in the photophysical properties of the complex. Such an effect might occur in systems incorporating identical metal centres with varying numbers of cyanide ligands attached to them, e.g. \([\text{Ru}(\text{CN})_3(\text{bpy})]^{2-}\) and \([\text{Ru}(\text{CN})_2(\text{bpy})_2]\) (§2.1.3), although to a lesser extent. The latter case might however suffer from spectral
congestion in ground and excited state TRIR spectroscopy in the v(CN) region, which might limit the insights which could be gained from spectroscopic study.

\[
\begin{align*}
\text{Figure 5.2.2: (RuCN)\textsubscript{x}HAT complexes.}
\end{align*}
\]

In conclusion a number of supramolecular systems have been designed, studied and analysed in conjunction with the photophysical models presented. The solvent-based nature of the switching of photophysical properties of polynuclear complexes incorporating the strongly solvatochromic [Ru(CN)\textsubscript{4}(bpy)]\textsuperscript{2-} unit has been demonstrated using UV/vis, luminescence and TRIR spectroscopies and electrochemistry. The utility of the v(CN) bands of the [Ru(CN)\textsubscript{4}(bpy)]\textsuperscript{2-} unit as a sensitive reporter of ground and excited state electron density in FTIR and TRIR spectra and as an efficient metal ion binder has also been shown. Possible applications and future extensions to the work in this Thesis have been presented.
5.3. References


Chapter 6:

Experimental Procedures
Chapter 6: Experimental Procedures

6.1. Materials, Measurements and Analysis

Materials

Chemicals were purchased from Sigma-Aldrich at the highest available purity. CH$_3$CN and CH$_3$OH were distilled over calcium hydride under an atmosphere of argon.

All complexes were synthesised by Mr. T. L. Easun and co-workers under the supervision of Prof. M. D. Ward at the University of Sheffield.$^1$

All experiments were conducted at room temperature (23 ± 2 °C) and atmospheric pressure unless otherwise indicated.

Spectroscopic Measurements

FTIR Spectroscopy

FTIR spectra were recorded in standard CaF$_2$ solution cells (Specac) using a Nicolet Avatar 360 FTIR spectrometer, recording 128 scans at a typical resolution of 2 cm$^{-1}$ with two levels of zero filling. For measurements in D$_2$O solution a pathlength of 50-100 μm was normally used and in CH$_3$CN solution a pathlength of 0.5 mm was normally used. During the collection of spectra in the 1500-1900 cm$^{-1}$ spectral region it was necessary to purge the spectrometer sample compartment with nitrogen or dry air due to strong absorptions of atmospheric water vapour in this region until no further spectral changes were observed, typically for 30-60 minutes.

UV/visible Absorption Spectroscopy

UV/visible absorption spectra were obtained in quartz cuvettes using a UNICAM UV-2 Spectrophotometer, typically with a data spacing of 1 nm using cuvettes of either 1 mm or 10 mm pathlength, depending on the absorbance of sample and solvent in the spectral regions of interest.
**Luminescence Spectroscopy**

Emission measurements were performed on a combined fluorescence lifetime and steady state spectrometer (Edinburgh Instruments FLS920). Non-aqueous solutions were thoroughly degassed using the freeze-pump-thaw technique, in specially modified 10 mm x 10 mm quartz cuvettes. Aqueous solutions were degassed using the same technique with care taken to ensure that the solidifying water did not rapidly expand and shatter the measuring vessel. The optical density was adjusted to ca. 0.2 at the excitation wavelength, which corresponds to a concentration of ca. $1 \times 10^{-5}$ M. Steady state emission and excitation spectra were obtained with a xenon arc lamp as the excitation source and were corrected for detector sensitivity. The emission monochromator wavelength was stepped in 1 nm increments. Emission lifetime measurements were performed using time-correlated single-photon counting (TCSPC), with a nanosecond H$_2$ flash lamp (repetition rate 40 Hz and pulse width ca. 2 ns) as the excitation source. The uncertainties quoted for emission lifetimes have been taken from the uncertainty of the fitted decay traces from Levenberg-Marquardt Non-Linear Least Squares Fitting.$^2$

Quantum yields, $\Phi_{em}$ were calculated from the following equation, using a solution of [Ru(bpy)$_3$]Cl$_2$ in H$_2$O as an emission standard,

$$\phi_{em} = \phi_r \frac{I}{I_r} \frac{OD_r}{OD} \frac{n^2}{n_r^2},$$

Where:

$\Phi_r$ = quantum yield of emission standard (0.028 for [Ru(bpy)$_3$]$^{2+}$ in H$_2$O at 298 K$^3$),

$I$ = integrated emission intensity of the sample,

$I_r$ = integrated emission intensity of the emission standard,

$OD$ = optical density of the sample at the excitation wavelength,

$OD_r$ = optical density of the emission standard at the excitation wavelength,

$n$ = refractive index of the solvent used for the sample,

$n_r$ = refractive index of the solvent used for the emission standard.

The fractional uncertainty in emission quantum yield is estimated as $\pm 10\%$.
A second apparatus was used for faster TCSPC experiments, the Edinburgh Instruments Mini-Tau Lifetime Spectrometer based at the Department of Chemistry, University of Sheffield. This enabled the undertaking of TCSPC experiments with an excitation wavelength of 405 nm with time resolution of below 1 ns. All other aspects of the experiments were identical.

**Time-resolved Infrared (TRIR) Spectroscopy**

For TRIR measurements of transition metal complexes concentrations ranging from $1 \times 10^{-4}$ to $5 \times 10^{-3}$ M were typically used. The choice of pathlength was dependent of the concentration of the solute and of solvent absorption in the spectral region under study. Typically, the pathlengths were 50-60 μm for D$_2$O and 500-1000 μm for CH$_3$CN and CH$_2$Cl$_2$. It should be noted that apparent discrepancies between time constants obtained using luminescence and TRIR spectroscopy may be rationalised in terms of the different experimental conditions required for the experiments. Two likely contributing factors are worth noting. The much higher analyte concentrations required for TRIR experiments may lead to self-quenching which is an effect that has been observed previously in TRIR spectra. Another possible reason for this is the inferior degassing protocol that is used during the PIRATE TRIR experiments and both of these factors would result in shorter $^3$MLCT lifetimes under the conditions of TRIR measurements compared to luminescence measurements. It should also be noted that in most cases the TRIR kinetic traces shown are ‘single-pixel’ in nature, in that a single detector pixel’s ΔAbsorbance signal is plotted versus time. In all traces, ΔAbsorbance = 0 is defined as the X-axis.

**Picosecond TRIR Spectroscopy**

The picosecond TRIR experiments were carried out on the PIRATE apparatus at the Central Laser Facility of the CCLRC Rutherford Appleton Laboratory. This apparatus has been described in detail previously. Briefly part of the output from a 1 kHz, 800 nm, 150 fs, 2 mJ Ti-Sapphire oscillator/regenerative amplifier (Spectra Physics Tsunami/Spitfire) was used to pump a white light continuum seeded β-BaB$_2$O$_4$ (BBO) optical parametric amplifier (OPA). The signal and idler produced by this OPA were
difference frequency mixed in a type I AgGaS$_2$ crystal to generate tuneable mid-infrared pulses (ca. 150 cm$^{-1}$ FWHM, 1 μJ), which were split to give probe and reference pulses. Second harmonic generation of the residual 800 nm light provided 400 nm pump pulses which were used for all ps-TRIR experiments described in this Thesis. Both the pump and probe pulses were focused to a diameter of 200-300 μm and co-linearly overlapped in the sample. Changes in infrared absorption at various pump-probe time delays were recorded by normalising the outputs from a pair of 64-element MCT infrared linear array detectors each monitoring independently the probe and reference pulses on a shot-by-shot basis. Data were collected in pump-on/pump-off pairs in order to minimise the effect of long-term drift in the laser intensity. Data points are spaced approximately 4-5 cm$^{-1}$ apart, giving a spectral resolution of ca. 8-9 cm$^{-1}$. Optical effects due to pump and probe laser interference are present in spectra recorded earlier than 1 picosecond, which is taken as the instrument response time (or IRF).

In all TRIR experiments dry, degassed solutions were saturated with argon and flowed through a home-built flow system incorporating an IR cell (Harrick Scientific Corp.) and a Teflon peristaltic pump (Cole Palmer). In the case of D$_2$O solutions the static samples were sealed inside the IR cell. The cell was rapidly oscillated in the plane perpendicular to the direction of the laser beams in order to minimise the potential build-up of decomposition products on the windows. All the samples were checked for the formation of decomposition products by on-line FTIR (vide supra) after each experiment.

**Nanosecond TRIR Spectroscopy**

The nanosecond TRIR spectra shown in this Thesis were obtained using the PIRATE apparatus at the Central Laser Facility of the CCLRC Rutherford Appleton Laboratory using a Nd:YAG laser (Advanced Optical Technology ACE) as the excitation source ($\lambda = 355$ nm) and the detection system detailed above.$^6$
Figure 6.1: A schematic of the PIRATE apparatus. Adapted from Ref.4

The Nottingham step-scan FTIR apparatus used in the early stages of this project has been described in detail elsewhere.7 Briefly the apparatus comprises of a commercially available step-scan FTIR spectrometer (Nicolet Magna 860) equipped with a 100 MHz 12-bit digitizer and a 50 MHz MCT detector interfaced to a Nd:YAG laser (Spectra Physics GCR-12). Synchronization of the Nd:YAG laser with data collection was achieved using a pulse generator (Stanford DG535). A commercially available IR cell (Harrick, 1 to 3 mm pathlength) was used with a home built flow system to minimise the build-up of decomposition products on the cell windows.

Electrochemistry

Cyclic Voltammetry

Cyclic voltammetric studies were carried out in anhydrous CH$_3$CN solutions (ca. 1 x 10$^{-3}$ M, with [Bu$_4$][BF$_4$] as the supporting electrolyte) using a potentiostat (EG & G Model 362) and a digitiser (EG & G Condecon 310). Cyclic voltammetry was carried out under an atmosphere of argon using a three-electrode arrangement in a single compartment cell. A Pt wire working electrode, a Pt wire secondary electrode and a
Ag/AgCl reference electrode, chemically isolated from the test solution via a bridge tube containing electrolyte solution and fitted with a porous vycor frit, were used in the cell. Redox potentials are quoted vs. the ferrocenium-ferrocene (Fc/Fc⁺) couple, which was used as an internal reference (at ca. +0.45 V vs. Ag/AgCl reference electrode). A scan rate of 100 mVs⁻¹ was used. Compensation for internal resistance was not applied.

Spectroelectrochemistry

UV/vis/nIR spectroelectrochemical experiments were carried out at the Department of Chemistry, University of Sheffield by Mr T. L. Easun and co-workers with an optically transparent electrochemical (OTE) cell (modified quartz cuvette, optical pathlength 0.5 mm). A three-electrode configuration, consisting of a Pt gauze working electrode (in a PTFE sheath), a Pt wire secondary electrode and a Ag/AgCl reference electrode, chemically isolated from the test solution via a bridge tube containing electrolyte solution and terminated in a porous frit, was used in the cell.

The potential at the working electrode was controlled by a Princeton Applied Research 273A potentiostat. The UV/vis/nIR spectra were recorded using a Varian Cary 5000 UV/vis/nIR spectrophotometer. The spectrometer cavity was purged with dry N₂ gas and temperature control at the sample was achieved by flowing cooled N₂ gas across the surface of the cell.

FTIR spectroelectrochemical experiments were carried out with an optically transparent thin-layer electrochemical (OTTLE) cell. The cell consisted of a modified IR solution cell (Specac, optical pathlength ca. 200 µm) equipped with CaF₂ windows and a three-electrode configuration; Pt/Rh minigrid working and (unbridged) secondary electrodes and a twisted Ag wire pseudo-reference electrode. The cell was positioned so that the beam passed through the centre of the working electrode; all other electrodes were masked from the IR beam. Potentials were controlled by a potentiostat (EG&G Model 362) and spectra were recorded at ambient temperature on an FTIR spectrometer (Nicolet Avatar) that was purged with dry air.
For UV/vis/nIR and FTIR spectroelectrochemical experiments the metal complex in solution was either electrolysed at constant potential (UV/vis) or the potential was stepped across the redox event (FTIR). For the single potential step experiments a potential typically 100 mV more positive than $E_p^a$ for an oxidation experiment and 100 mV more negative than $E_p^c$ for a reduction was applied to the working electrode. The redox process was considered complete when consecutive spectra were identical. The chemical reversibility of the process was investigated by applying a potential at the working electrode sufficient to re-oxidise or re-reduce the electrogenerated product. These potentials were typically 100 mV more negative than $E_p^c$ to reverse an oxidation process or 100 mV more positive than $E_p^a$ to reverse a reduction process. The process was considered to be reversible under the conditions of the experiment if the spectroscopic profile of the starting material was reproduced. The kinetics of the redox process may not allow full return to the starting material within the timescale of the experiment. In such cases isosbestic points were taken to indicate reversibility of the spectral profile.

**Analysis of TRIR Data**

The basic methodology employed for the initial analysis of the raw data, generated in the ps-TRIR experiments in the PIRATE facility at the Central Laser Facility of the Rutherford Appleton Laboratory has been described previously.\(^\text{10}\)

In each experiment up to 100 pump-probe time delays could be measured. The experiment could be accumulated for any number of cycles and for each cycle the computer randomised the order in which the time delays were measured. The acquisition time at each time delay was determined by the user prior to the experiment on the basis of the signal-to-noise ratio in the spectra. This in turn was determined by the signal strength (sample concentration) and solvent absorption in the spectral region of interest. For ps-TRIR spectra for CH$_3$CN or CH$_2$Cl$_2$ solutions of metal carbonyl or cyanide complexes in the metal carbonyl region, 20 time delays measured over 2 cycles with 15 s data acquisition time at each time delay was usually sufficient to provide a good signal-to-noise ratio. For ps-TRIR spectra of solutions of transition
metal complexes in D$_2$O typically 15-25 time delays measured over 4-6 cycles with 15-30 s data acquisition time at each time delay was used.

Custom software (EMToF and FeTian, written by Dr. Xue-Zhong Sun, University of Nottingham) was used to automatically generate the “ΔAbsorbance Difference” TRIR spectra from the .dao data files for each time delay. The spectra contain 64 data points, representing the 64 pixels of the linear IR array detectors. Calibration of these spectra from pixel numbers into wavenumbers was achieved by curve-fitting one of the pixel spectra and comparing the positions of the bleach bands (in pixels) with the ground state FTIR bands (in cm$^{-1}$) taking into account a spectrograph dispersion of ca. 12 nm/pixel. This procedure was complicated somewhat by the presence of “dead pixels” in the detector array, most notably pixel number 39. This resulted in some erroneous signals near the centre of the spectra collected which were disregarded when data was analysed.

Curve-fitting was performed using commercially available software (OriginPro 7e, OriginLab Inc.). Individual band parameters for the transient and parent bands in the ps-TRIR spectra were obtained from multi-curve fitting of the ps-TRIR spectra presuming their Lorentzian band shape in non-polar solvents, with the facility to employ Gaussian band shapes in case of line broadening in polar solvents. For ps-TRIR spectra in the metal carbonyl region the Lorentzian parameters of the parent bands (band positions, widths and the ratio between the parent bands) were obtained by fitting the ground state FTIR spectrum and kept constant during the fitting procedure. The parameters of vibrationally cool bands were determined by curve-fitting of one of the spectra at late time delays, typically 100 ps, and again were fixed as constants during the fitting procedure.

In many ν(CN) TRIR spectra presented in this work, the parent bands are heavily overlapped with the transient bands and this complicates the calibration of the ps-TRIR spectra. In extreme cases a ground state IR spectrum was recorded with ps-TRIR apparatus by plotting the log$_{10}$ of the ratio between reference only signals for the sample and pure solvent to allow a precise calibration and fitting of time-resolved
spectra. The obtained ground state IR spectrum was again fitted with the sum of Lorentzian bands and the band parameters were kept constant during the fit of ps-TRIR spectra.

As a result of heavy ground and excited state band overlap in the ps and ns-TRIR spectra in this Thesis it was often not feasible to obtain kinetic information from the multi-curve fitting of the ps-TRIR spectra due to the large number of overlapping bands and associated uncertainty in the band parameters. In this case the kinetic information was obtained by recording intensity of the signal in a single point, at the maximum of the band of interest or in the point where this band is least likely to be influenced by overlap with surrounding bands. When this was the case the Figure captions mention this explicitly.

**Conformational Calculations using SPARTAN**

Conformational analysis was conducted using the commercially available software package SPARTAN (Wavefunction Inc.) by constructing the supramolecular system and performing geometric optimisation through semi-empirical energy minimisation using the PM3 algorithm. Following this the Conformational Analysis function was performed which in combination with manual verification gave approximate internuclear distances between metal centres. These are meant to be used as a “rule of thumb” and molecular dynamics or full DFT calculations with a supercomputer would be preferable to obtain more precise and meaningful data.
6.2. References