

Lower-rim ferrocenyl substituted calixarenes : new electrochemical sensors for anions

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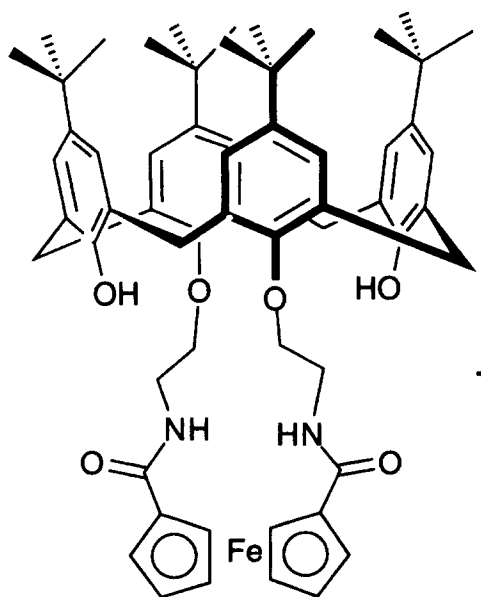
(Received 23 July 1997; accepted 16 September 1997)

Abstract—New ferrocene substituted calix[4 and 5]arenes have been prepared and the crystal structure of a lower-rim substituted bis ferrocene calix[4]arene (**7**) has been elucidated. The respective ferrocene/ferrocenium redox-couples of compounds **6** (a calix[4]arene tetra ferrocene amide) and **8** (a calix[5]arene pentaferrocene amide) are shown to be significantly cathodically perturbed in the presence of anions by up to 160 mV in the presence of dihydrogen phosphate. © 1998 Elsevier Science Ltd. All rights reserved

Keywords: sensors ; ferrocene ; calixarene ; anion binding ; amide ; electrochemistry.

Anions are ubiquitous in biological systems [1]. Many proteins contain inorganic phosphate or sulfate anions coordinated via hydrogen bonds from water, amino-acid residues or amide bonds [2]. Additionally excess phosphate pollutants in the environment cause eutrophication of rivers and inland waterways [3], whereas nitrate pollutants and their metabolites are believed to be carcinogenic [4]. There is therefore great interest in the synthesis of receptor species containing hydrogen bond donor arrays [5,6], Lewis acid centres [7] or positive charges [8,9] as sequestering agents or sensors for anions. Early work by ourselves [10–13] and others [14,15] established that amide N—H groups in acyclic synthetic receptor molecules can coordinate anions. As part of this work we demonstrated that the calix[4]arene **1**, which contains a lower-rim ferrocene amide bridge is capable of electrochemically detecting the presence of anions *via* perturbations of the ferrocene/ferrocenium redox couple [16].

As a further extension to this work we decided to synthesise a series of calixarenes containing multiple amidoferrocenyl residues at the lower rim and study these systems as potential amperometric sensors for anions.



RESULTS AND DISCUSSION

Syntheses

p-tert-Butylcalix[4]arene tetra (acid chloride) **2**, *p*-tert-butylcalix[4]arene bis (acid chloride) **3**, *p*-tert-

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butyralix[5]arene penta (acid chloride) **4** [17,18] and ferrocenemethylamine **5** [19] were synthesised *via* literature procedures. An analogous synthetic procedure was employed in the preparation of each of the new calixarene-ferrocene conjugate compounds. Typically, the calixarene acid chloride was dissolved in dry dichloromethane, a 25% excess of ferrocenemethylamine **5** and triethylamine added, and the reaction mixture stirred under nitrogen overnight, in darkness. After washing with water, drying over magnesium sulfate and removal of solvent *in vacuo*, a pale brown solid was generally produced. This was purified by column chromatography (silica support) using dichloromethane/methanol as eluent affording the *p*-tert-butylcalix[4]arene tetra amidoferrocene **6** in 21% yield, the *p*-tert-butylcalix[4]arene bis amidoferrocene **7** in 29% yield and the *p*-tert-butylcalix[4]arene bis amidoferrocene **8** in 16% yield (Scheme 1). The moderate yields of these reactions may be due to the incomplete reaction of the calixarene acid chlorides with ferrocenemethylamine (due to steric interactions), leading to incomplete substitution of ferrocene at the lower rim of the calixarene.

Crystallography

Diffraction quality single crystals of **7** were grown from a dichloromethane/methanol solvent mixture. The structure confirms that **7** adopts the cone conformation in the solid state (Fig. 1) with two ferrocene groups appended at the lower rim. The calixarene has usual C₂ distortion with the angles between the plane of the four methylene groups and the four phenyl rings being 44.2, 66.0, 52.9 and 70.6°, respectively. As can be seen in the figure, there is a dichloromethane molecule encapsulated within the cavity. Crystal data is given in Table 1.

Anion coordination studies: NMR titrations

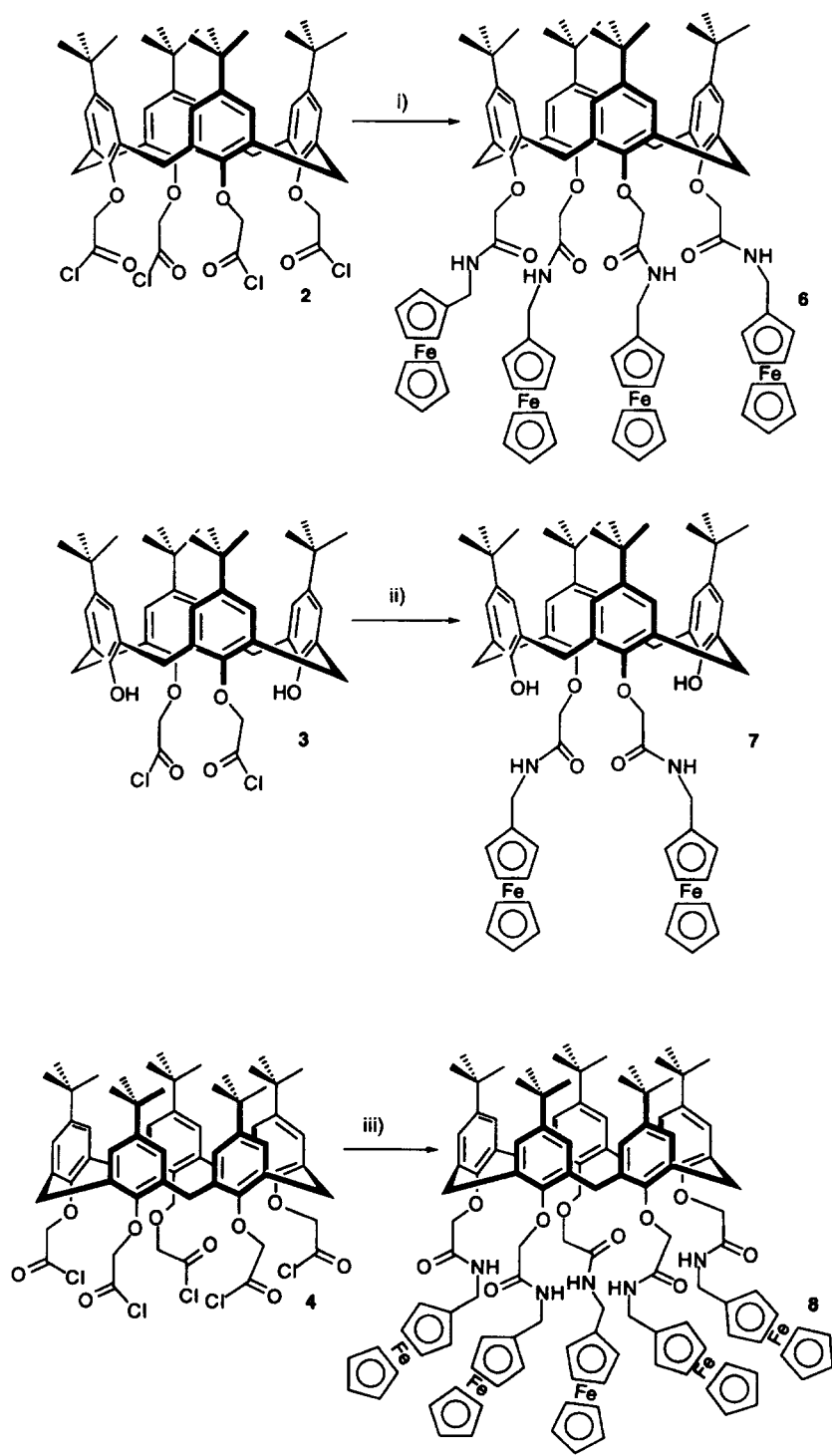
¹H NMR titrations were carried out in dichloromethane-*d*₂ solutions of compounds **6**, **7** and **8**. In a typical titration, 5 × 10⁻⁶ mol of receptor was dissolved in 0.5 ml dichloromethane-*d*₂ and equivalents of the anion added as *n*-Bu₄N⁺X⁻ (X⁻ = Cl⁻, HSO₄⁻ or H₂PO₄⁻) in dichloromethane-*d*₂ (0.1 mol dm⁻³) solution. The shifts of resonances of protons involved in anion coordination were then recorded and plotted as a function of the amount of anion added. Disappointingly, in the case of **7**, no significant shifts, i.e. greater than 0.01 ppm, were observed indicating that there is little interaction between this molecule and the anionic species. Partial substitution reduces the number of amide hydrogen bond donor sites at the lower rim of **7** compared to **6** and **8**, and this may be the cause for the poor anion coordination ability of this ligand. Another possibility is that in solution, the phenolic OH groups are hydrogen bonded to the amide functionalities via either

OH...O=C or HO...H—N hydrogen bonds so blocking the potential coordinating sites. This is reflected in the chemical shift of the amide resonance in the ¹H NMR spectrum of **7** which occurs at 8.74 ppm (i.e. it is deshielded relative to the amide protons of **6** and **8** which resonate at 7.57 and 7.42 ppm, respectively). The tetra- and penta- substituted derivatives, however, display significant downfield shifts of the amide NH proton (approximately 0.1 ppm) in the presence of anions (Figs 2 and 3). This indicates that the amide NH protons are coordinating to the added anions. Titrations with compounds **6** or **8** with chloride anions produced titration curves which could be fitted to a 1 : 1 calixarene/anion stoichiometric model. Stability constants were determined with the EQNMR computer program [20] and are shown in Table 2. Compound **6** forms a more stable complex with chloride (55 M⁻¹) than does compound **8** (15 M⁻¹). This may be due to a spacial fit effect, with chloride coordinating to the amide moieties of **6** more efficiently than to those of the calix[5]arene **8**. An adequate fit could not be obtained for the titrations of **6** and **8** against hydrogen sulfate and dihydrogen phosphate to models consistent with 1 : 1 receptor/anion stoichiometries. This may be due to the formation of complexes of other stoichiometries, or aggregation of the anions in the relatively non-polar dichloromethane solvent. Unfortunately, precipitation problems were encountered during the titration studies upon addition of higher concentrations of anions. This precluded the limiting chemical shifts of the anion complexes in these experiments from being found.

Anion coordination studies: electrochemistry

The electrochemical properties of **6**, **7** and **8** were investigated using cyclic voltammetry [21] in a dichloromethane/acetonitrile (35 : 65 v/v) solvent mixture, with (*n*-Bu)₄NBF₄ as the supporting electrolyte. The working electrode used was a 5 mm glassy carbon disk and the counter electrode consisted of a platinum mesh. A Ag/Ag⁺ electrode was used as reference (+330 mV vs S.C.E.). Argon gas saturated with dichloromethane was continually bubbled through the cell to de-oxygenate the solution. It must be noted that in all cases electrochemical experimentation proved difficult with many precipitation and solubility problems encountered, as well as frequent adsorption of the oxidised species onto the electrode in the presence of anions. Each of the ligands exhibited a single quasi-reversible oxidation wave in the region 0.12 to 0.20 V (vs Ag/Ag⁺) (Table 2, Fig. 4) suggesting the ferrocene moieties are behaving independently of one another and are therefore being oxidised at the same potential.

Thin layer coulometry was used to confirm that **6**, **7**, and **8** are undergoing four, two and five electron oxidations, respectively. In each case, the number of electrons transferred to the compound in question was found to be approximately equal to the number of ferrocene subunits present.



- i) 5.0 FcCH₂NH₂, 5.0 Et₃N / CH₂Cl₂
ii) 2.5 FcCH₂NH₂, 2.5 Et₃N / CH₂Cl₂
iii) 6.0 FcCH₂NH₂, 6.0 Et₃N / CH₂Cl₂

Scheme 1.

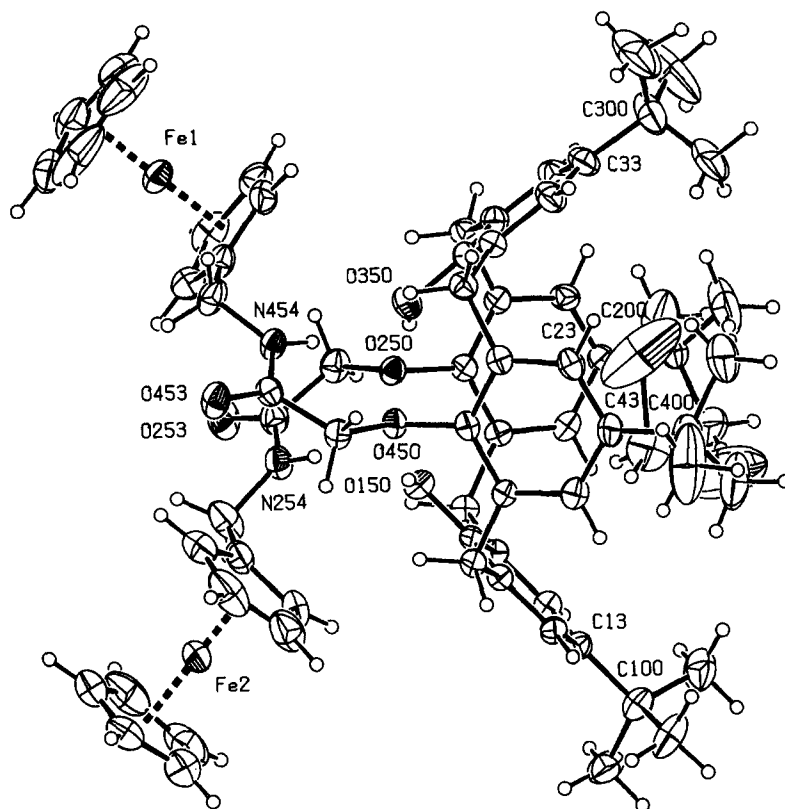


Fig. 1. The structure of **7**, with thermal ellipsoids drawn at the 30% probability level. Hydrogen atoms are included as small circles of arbitrary small radius. Intramolecular hydrogen bonds (not shown) are formed between the amide nitrogen atoms and the phenolic OH groups {N(254)···O(150) 2.98 Å, N(454)···O(350) 2.94 Å}.

Cyclic voltammograms were also recorded after progressive addition of stoichiometric equivalents of anionic guests as 0.1 M solutions of (*n*-Bu₄N)H₂PO₄, (*n*-Bu₄N)HSO₄ and (*n*-Bu₄N)Cl in CH₂Cl₂/CH₃CN (35:65 v/v) to the electrochemical solutions of **6**, **7** and **8** (1×10^{-3} M) in CH₂Cl₂/CH₃CN (35:65 v/v), and the results are summarised in Table 3.

The addition of anions to electrochemical solutions of **6** and **8** produced cathodic shifts of the ferrocene/ferrocenium redox couple with the relative magnitudes following the order H₂PO₄⁻ > Cl⁻ > HSO₄⁻. The most significant results were observed with compounds **6** and **8** with H₂PO₄⁻ anions, Fig. 5. Similar results were observed with compound **1** [16].

The addition of dihydrogen phosphate anions causes a 160 mV cathodic shift in the ferrocene/ferrocenium redox couple. This large cathodic shift indicates a strong stabilisation of the oxidised forms of **6** and **8**, in the presence of dihydrogen phosphate. This may be due to dihydrogen phosphate coordination to the amide moieties of the calixarene, a consequence of suitable cavity size as defined by the pendant ferrocene amide groups. This places the anion in close proximity to the ferrocene/ferrocenium moiety,

so perturbing the redox potential of these groups. No analogous result was observed with **7** (in agreement with the prior ¹H NMR titration studies).

Conclusions

In conclusion the redox-active ferroceneamide derivatised calix[4,5]arenes (**6**) and (**8**) have been shown to bind and electrochemically detect anions. The stability constants of **6** and **8** with chloride anions have been determined and found to be 55 M⁻¹ and 15 M⁻¹, respectively. Compounds **6** and **8** electrochemically recognise Cl⁻, HSO₄⁻ and H₂PO₄⁻ anions with H₂PO₄⁻ producing the largest shift of the ferrocene/ferrocenium redox couple (160 mV).

EXPERIMENTAL

General methods

All elemental analyses were carried out by the Inorganic Chemistry Laboratory Microanalysis Service. NMR spectra were recorded on a Bruker AM300 NMR spectrometer. Infrared spectra were recorded

Table 1. Crystal data and structure refinement for **7**

Empirical formula	C ₇₃ H ₉₂ Cl ₂ Fe ₂ N ₂ O ₈
Formula weight	1308.09
Temperature	293(2) K
Wavelength	0.71070 Å
Crystal system	Triclinic
Space group	<i>P</i> $\bar{1}$
Unit cell dimensions	<i>a</i> = 13.772(10) Å <i>b</i> = 16.550(10) Å <i>c</i> = 17.097(10) Å <i>a</i> = 66.883(1)° <i>b</i> = 85.546(1)° <i>γ</i> = 77.413(1)°
Volume	3498(4) Å ³
<i>Z</i>	2
Density (calculated)	1.242 Mg/m ³
Absorption coefficient	0.55 mm ⁻¹
Crystal size	0.3 × 0.3 × 0.25 mm
Theta range for data collection	2.22 to 24.98°
Index ranges	0 ≤ <i>h</i> ≤ 14, −18 ≤ <i>k</i> ≤ 19, −20 ≤ <i>l</i> ≤ 20
Independent reflections	10,808
Refinement method	Full-matrix least-squares on <i>F</i> ²
Data/parameters	10,808/805
Goodness-of-fit on <i>F</i> ²	0.831
Final <i>R</i> indices [<i>I</i> > 2σ(<i>I</i>)]	<i>R</i> 1 = 0.075, <i>wR</i> 2 = 0.232
<i>R</i> indices (all data)	<i>R</i> 1 = 0.112, <i>wR</i> 2 = 0.304
Largest diff. peak and hole	1.006 and −1.396 e.Å ⁻³

$$R_1 = \frac{\sum |F_o| - |F_c|}{\sum |F_o|} \quad wR_2 = \frac{[\sum w(F_o^2 - F_c^2)] / \sum w(F_o^2)]^{1/2}}$$

on a Mattson "Polaris" and Perkin Elmer "1710" Infrared Fourier Transform Spectrometers. Electrochemical experiments were conducted on a Princeton Applied Research Potentiostat/Galvanostat Model 273. Fast atom bombardment (FAB) mass spectra were carried out at the University of Wales, Swansea by the SERC/EPSCRC Mass Spectrometry Service. Melting points were recorded in open capillaries on a Gallenkamp melting point apparatus and are uncorrected.

Synthesis of *p*-tert-butylcalix[4]arenetetraferrocene-amide (**6**)

Ferrocenemethylamine (0.61 g, 2.84 mmol) was dissolved in dry CH₂Cl₂ (30 ml) and added to *p*-tert-butylcalix[4]arene bis (acid chloride) (0.56 mmol). Et₃N (0.29 g, 2.87 mmol) was then added and the reaction mixture stirred under nitrogen overnight. After washing with water (4 × 25 ml), the organic layer was dried over magnesium sulfate. The solvent was then removed under reduced pressure to give a pale brown solid which was purified by column chromatography on silica gel, gradient eluting with CH₂Cl₂/MeOH (gradient elution: concentration of

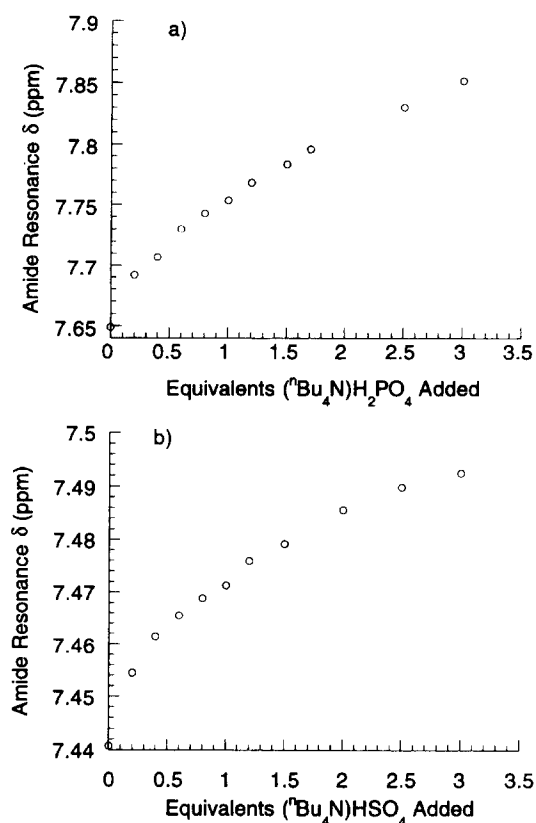


Fig. 2. NMR titrations (CD₂Cl₂) (a) **6** with (*n*-Bu₄N)H₂PO₄ and (b) **8** with *n*-Bu₄NHSO₄.

methanol from 0% to 5%). The product was collected, reduced *in vacuo* and then precipitated as a yellow powder from a dichloromethane/methanol solvent mixture as a yellow powder (0.2 g, 21%). ¹H NMR (CDCl₃, 300 MHz) δ: 1.05 (s, 36H, C(CH₃)₃), 3.12 (d, ²J = 12.5 Hz, 4H, ArCH₂Ar:H_{eq}), 4.11 (m, 8H, FcH), 4.16 (s, 20H, FcH), 4.25 (m, 8H, FcH), 4.32 (d, 8H, —NHCH₂Fc) and (d, 4H, ArCH₂Ar:H_{ax}) coincident, 4.36 (s, 8H, OCH₂), 5.31 (CH₂Cl₂), 6.72 (s, 8H, ArH), 7.57 (t, 4H, NH). ¹³C NMR/DEPT (CDCl₃, 75.42 MHz) δ: 31.33 ((CH₃)₃C), 31.49 (ArCH₂Ar), 33.66 (C(CH₃)₃), 38.70 (—OCH₂C(O)—), (68.07, 68.62, 68.63, (FcH)), 74.65 (NHCH₂Fc), 84.93 (Fc), 125.76 (ArH), (132.72, 145.71, 152.77 (Ar)), 169.02 (C=O). Microanalysis Calc. for C₉₆H₁₀₈O₈N₄Fe₄ · 1/3(CH₂Cl₂) C, 68.16; H, 6.45; N, 3.30. Found: C, 68.52; H, 6.53; N, 3.48. FAB MS: MNa⁺@1693. Infrared Data (KBr Disc) 1662 cm⁻¹ (carbonyl stretch) m.p. 255°C.

Synthesis of *p*-tert-butylcalix[4]arenebisferrocene-amide (**7**)

Ferrocenemethylamine (0.29 g, 1.35 mmol) and Et₃N (0.14 g, 1.38 mmol) were dissolved in dry dichloromethane (25 ml) and added to a round bottomed

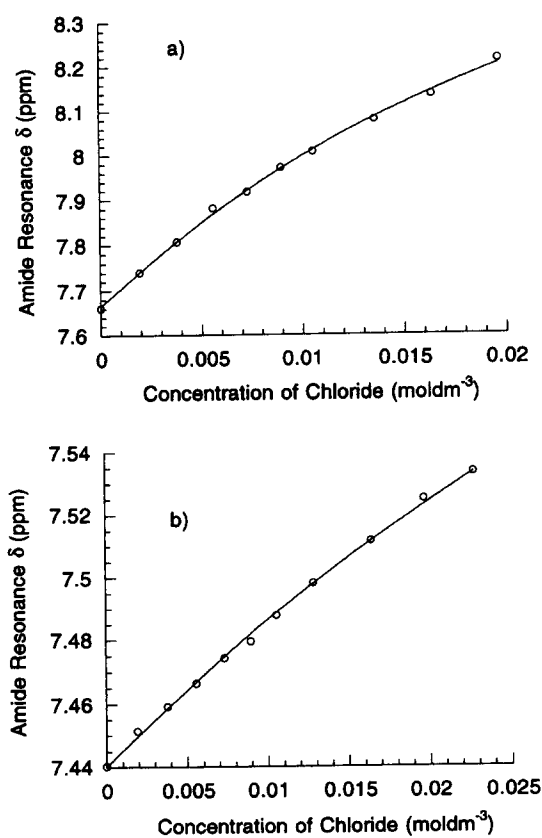


Fig. 3. NMR titrations (CD_2Cl_2) (a) **6** with Cl^- and (b) **8** with Cl^- .

Table 2. Stability constants obtained by ^1H NMR titrations with chloride in dichloromethane- d_2 (maximum error $\pm 20\%$)

Ligand	Stability constant (M^{-1})
6	55
8	15

flask containing *p*-tert-butylcalix[4]arene bis (acid chloride) (0.54 mmol). The reaction mixture was then left stirring under nitrogen, in the dark, for twenty hours to give a deep red solution. This was washed with water (4×25 ml), and the solvent removed *in vacuo* to give orange/red crystals. Purification by column chromatography on silica, gradient eluting with dichloromethane/methanol (gradient elution: concentration of methanol from 0% to 5%). The product was collected, reduced *in vacuo* to give orange/yellow crystals (0.18 g, 29%). ^1H NMR (CDCl_3 , 300 MHz) δ : 0.95 (s, 18H, $(\text{CH}_3)_3\text{C}$), 1.26 (s, 18H, $(\text{CH}_3)_3\text{C}$), 3.27 (d, $^2J = 13.3$ Hz, 4H, $\text{ArCH}_2\text{Ar} : \text{H}_{\text{eq}}$), 3.87 (d,

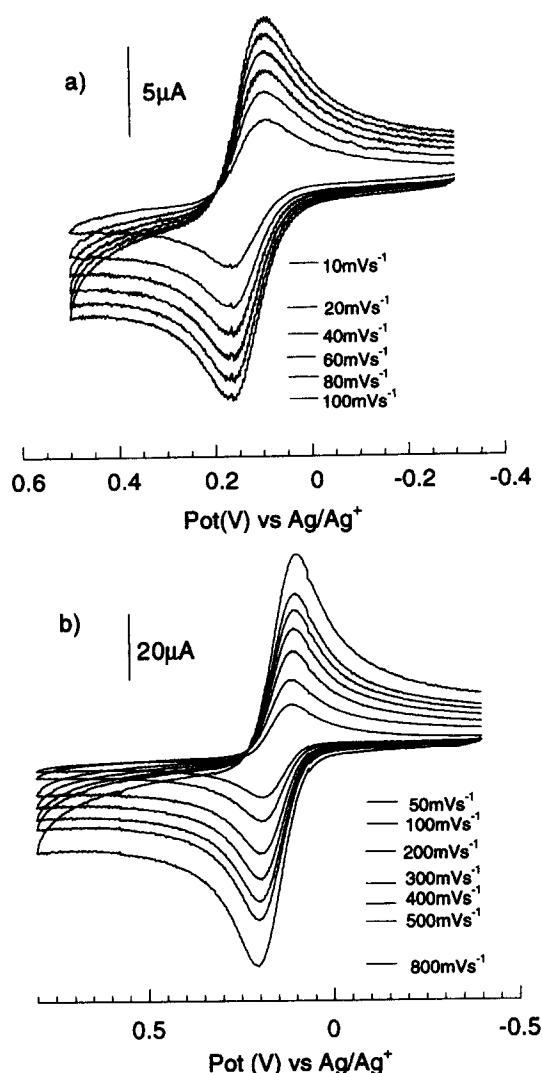


Fig. 4. Quasi-reversible cyclic voltammograms of (a) (**6**) and (b) (**7**) recorded at 1.0×10^{-3} mol dm^{-3} (ligand). Glassy carbon working electrode, Ag^+/Ag reference electrode ($+330$ mV vs S.C.E.). Recorded in $\text{CH}_2\text{Cl}_2/\text{CH}_3\text{CN}$ (35:65 v/v).

$^2J = 13.3$ Hz, 4H, $\text{ArCH}_2\text{Ar} : \text{H}_{\text{ax}}$), 4.10 (s, 4H, FcH), 4.13 (m, 4H, FcH), 4.15 (m, 4H, FcH), 4.28 (d, 4H, NHCH_2Fc), 4.38 (s, 4H, $(\text{OCH}_2\text{C}(\text{O}))$), 5.31 (CH_2Cl_2), 6.78 (s, 4H, ArH), 6.99 (s, 4H, ArH), 8.74 (t, 2H, NH). ^{13}C NMR/DEPT (CDCl_3 , 75.42 MHz) δ : 30.90 ($(\text{CH}_3)_3\text{C}$), 31.83 ($(\text{CH}_3)_3\text{C}$), 31.88 (ArCH_2Ar), 33.89 ($(\text{CH}_3)_3\text{C}$), 34.20 ($(\text{CH}_3)_3\text{C}$), 39.91 (OCH_2), [68.34, 68.89, 69.56 (FcH)], 74.77 (NHCH_2Fc), 82.47 (Fc), 125.32 (ArH), 126.00 (ArH), [126.73, 132.10, 142.71, 148.21, 149.23 (Ar)], 167.76 ($\text{C}=\text{O}$). Microanalysis: Calc. for $\text{C}_{70}\text{H}_{82}\text{O}_6\text{N}_2\text{Fe}_2 \cdot 1/3(\text{CH}_2\text{Cl}_2)$ C, 71.14; H, 7.02; N, 2.36. Found: C, 71.88; H, 7.18; N, 2.48. FAB MS: $\text{M}^+ @ 1158$; $\text{MNa}^+ @ 1181$. Infrared Data (KBr Disk): 1681 cm^{-1} (carbonyl stretch). m.p. 245°C (dec).

Table 3. Maximum cathodic shifts in the presence of excess anions

	Ligand		
	6	7	8
$E_{1/2}/V(\text{vs Ag}/\text{Ag}^+)$	0.1242	0.1892	0.1318
$\Delta(\text{Cl}^-)/\text{mV}$	15	<5	20
$\Delta E(\text{HSO}_4^-)/\text{mV}$	15	<5	20
$\Delta E(\text{H}_2\text{PO}_4^-)/\text{mV}$	160	<5	160

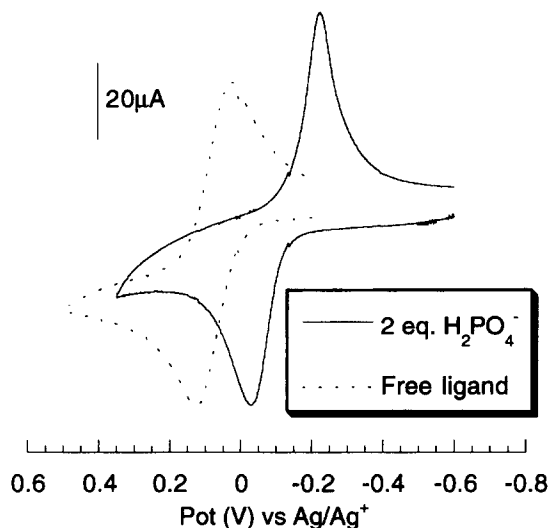


Fig. 5. Cyclic voltammograms of compound **6** in the absence and presence of two equivalents of dihydrogen phosphate in 35% $\text{CH}_2\text{Cl}_2/\text{CH}_3\text{CN}$. Scan rate 100 mV s^{-1} , glassy carbon working electrode, Ag^+/Ag reference electrode (+330 mV vs S.C.E.). Recorded in $\text{CH}_2\text{Cl}_2/\text{CH}_3\text{CN}$ (35:65 v/v).

Synthesis of *p*-*tert*-butylcalix[5]arenepentaferroceneamide (**8**)

Ferrocenemethylamine (0.60 g, 2.79 mmol), triethylamine (0.28 g, 2.77 mmol) and *p*-*tert*-butylcalix[5]arene penta (acid chloride) (4.7 mmol) were dissolved in dry dichloromethane (30 ml) and the reaction mixture stirred for eighteen hours. After washing with water, the product was dried over magnesium sulfate and the solvent removed under reduced pressure to give brown/orange crystals. These were purified by column chromatography on silica gel, gradient eluting with dichloromethane/methanol. The product was collected, reduced *in vacuo*, and then precipitated as a yellow powder (0.16 g, 16%) from a dichloromethane/methanol mixture (gradient elution: concentration of methanol from 0% to 5%). ^1H NMR (CDCl_3 , 300 MHz) δ : 1.02 (s, 45H, $(\text{CH}_3)_3\text{C}$), 3.31 (d, 5H, $J = 14.0 \text{ Hz}$, $\text{ArCH}_2\text{Ar}:\text{H}_{\text{ax}}$), 4.12 (s, 10H, OCH_2), 4.17 (s, 25H, FcH), 4.24 (m, 10H, FcH), 4.26 (m, 10H, FcH), 4.27 (d, 10H, $-\text{NHCH}_2$), 4.47 (d,

5H, $J = 14.0 \text{ Hz}$, $\text{ArCH}_2\text{Ar}:\text{H}_{\text{ax}}$), 6.88 (s, 10H, ArH), 7.42 (t, 5H, NH). ^{13}C NMR (CDCl_3 , 75.42 MHz) δ : 30.02 (ArCH_2Ar), 31.30 ($(\text{CH}_3)_3\text{C}$), 34.05 ($(\text{CH}_3)_3\text{C}$), 38.83 (OCH_2CO), [68.14, 68.59, 68.73 (FcH)], 72.93 (NHCH_2), 84.93 (Fc), [126.35, 132.62, 146.11, 151.15 (Ar)], 169.01 (C=O). Microanalysis: calc. for $\text{C}_{120}\text{H}_{135}\text{O}_{10}\text{N}_5\text{Fe}_5$: C, 69.07; H, 6.52; N, 3.36 Found: C, 69.41; H, 6.75; N, 3.33. FAB MS: MH^+ @2087, MNa^+ @2109. Infrared Data (KBr Disk): 1672 cm^{-1} (carbonyl stretch). m.p. 295°C (dec).

Crystallography

Single crystal X-ray analysis of **7** was conducted using $\text{MoK}\alpha$ radiation with the MARresearch Image Plate System. The crystal was positioned at 75 mm from the X-ray source. 90 frames were measured at 2' intervals with a counting time of 2 mins. Data analysis was carried out with the XDS program [22]. The structure was solved using direct methods with the Shelx86 program [23]. Apart from the substituted calixarene, there were several solvent molecules and two methanol and one dichloromethane molecules were located, though there were indications of other disordered solvent molecules. The non-hydrogen atoms were refined with anisotropic thermal parameters apart from the carbon and oxygen atoms in the solvent. All hydrogen atoms were included in calculated positions apart from those in the solvent molecules. The structure was then refined using Shelx [24], R-factors are given in Table 1. All calculations were carried out on a Silicon Graphics R4000 Workstation at the University of Reading.

Acknowledgements—We thank the EPSRC for a studentship (PAG) and the EPSRC together with the University of Reading for the provision of the image plate diffractometer.

REFERENCES

1. Stryer, L., *Biochemistry*, 4th edn ed. W. H. Freeman and Co., New York, 1988.
2. Chakrabarti, P., *J. Mol. Biol.*, 1993, **234**, 463–482.
3. Moss, B., *Chem. and Ind.*, 1996, 407.
4. Glidewell, C., *Chem. Br.*, 1990, **26**, No. 2, 137.
5. Gale, P. A., Sessler, J. L., Král, V. and Lynch, V., *J. Am. Chem. Soc.*, 1996, **118**, 5140.
6. Gale, P. A., Sessler, J. L., Allen, W. E., Tvermoes, N. A. and Lynch, V., *Chem. Commun.*, 1997, 665.
7. Beer, P. D., *Chem. Commun.*, 1996, 689.
8. Sessler, J. L., Cyr, M. J., Lynch, V., McGhee, E. and Ibers, J. A., *J. Am. Chem. Soc.*, 1990, **112**, 2810.
9. Sessler, J. L., Cyr, M., Furuta, H., Král, V., Mody, T., Morishima, T., Shionoya, M. and Weghorn, S. J., *Pure and Appl. Chem.*, 1993, **65**, 393.
10. Beer, P. D., Heseck, D., Hodacova, J. and Stokes, S. E., *J. Chem. Soc. Chem. Commun.*, 1992, 270.
11. Beer, P. D., Hazlewood, C., Heseck, D., Hoda-

- cova, J. and Stokes, S. E., *J. Chem. Soc. Dalton Trans.*, 1993, 1327.
12. Beer, P. D., Drew, M. G. B., Hazlewood, C., Heseck, D. and Stokes, S. E., *J. Chem. Soc. Chem. Commun.*, 1993, 229.
 13. Beer, P. D., Drew, M. G. B., Hazlewood, C., Heseck, D., Hodacova, J. and Stokes, S. E., *J. Chem. Soc. Chem. Commun.*, 1993, 229.
 14. Pascal, R. A., Spergel, J. and van Engen, D. V., *Tetrahedron Lett.*, 1986, **27**, 4099.
 15. Valiyaveetil, S., Engbersen, J. F. J., Verboom, W. and Reinhoudt, D. N., *Angew. Chem. Int. Ed. Engl.*, 1993, **32**, 900.
 16. Beer, P. D., Chen, Z., Goulden, A. J., Graydon, A., Stokes, S. E. and Wear, T., *J. Chem. Soc. Chem. Commun.*, 1993, 1834.
 17. McKervey, M. A., Seward, E. M., Ferguson, G., Ruhl, B. and Harris, S. J., *J. Chem. Soc. Chem. Commun.*, 1985, 388.
 18. Arnaud-Neu, F., Schwing-Weill, M. J., Ziat, K., Cremin, S., Harris, S. J. and McKervey, M. A., *New J. Chem.*, 1991, **15**, 33.
 19. Spencer, P., Polyaza redox-active acyclic and macrocyclic compounds designed to bind cations and anions, D.Phil. thesis, University of Oxford, 1994.
 20. Hynes, M. J., *J. Chem. Soc., Dalton Trans.*, 1993, 311.
 21. *Instrumental Methods in Electrochemistry*, Southampton Electrochemistry Group, Ellis Horwood Limited, University of Southampton, 1990.
 22. Kabsch, W., *J. Appl. Cryst.*, 1988, **21**, 916.
 23. Sheldrick, G. M., *Acta Cryst.*, 1990, **A46**, 467.
 24. Sheldrick, G. M., Pers. comm., 1993.