1 **Hydrothermal Decomposition of Amino Acids and Origins of Prebiotic Meteoritic Organic Compounds**

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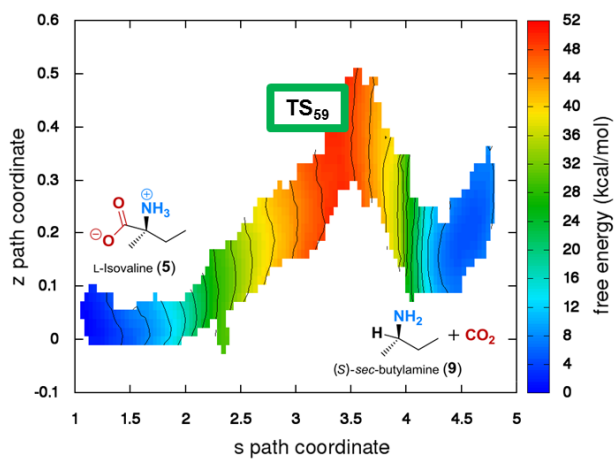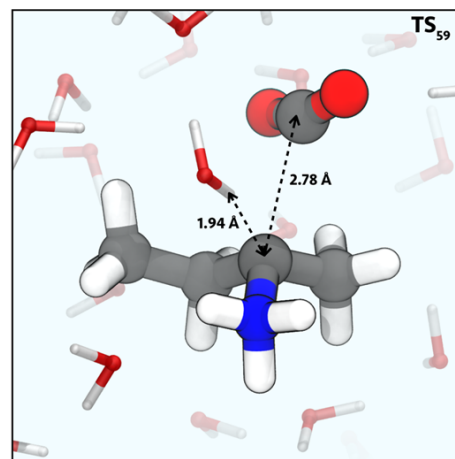
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ABSTRACT

The organic compounds found in carbonaceous chondrite meteorites provide insight into primordial solar system chemistry. Evaluating the formation and decomposition mechanisms of meteoritic amino acids may aid our understanding of the origins of life and homochirality on Earth. The amino acid glycine is widespread in meteorites and other extraterrestrial environments; other amino acids, such as isovaline, are found with enantiomeric excesses in some meteorites. The relationship between meteoritic amino acids and other compounds with similar molecular structures, such as aliphatic monoamines and monocarboxylic acids is unclear; experimental results evaluating the decomposition of amino acids have produced inconclusive results about the preferred pathways, reaction intermediates, and if the conditions applied may be compatible with those occurring inside meteoritic parent bodies. In this work, we performed extensive tandem metadynamics, umbrella sampling, and committor analysis to simulate the neutral mild hydrothermal decomposition mechanisms of glycine and isovaline and put them into context for the origins of meteoritic organic compounds. Our *ab initio* simulations aimed to determine free energy profiles and decomposition pathways for glycine and isovaline. We found that under our modeled conditions, methylammonium, glycolic acid, and *sec*-butylamine are the most likely decomposition products. These results suggest that meteoritic aliphatic monocarboxylic acids are not produced from decomposition of meteoritic amino acids. Our results also indicate that the decomposition of L-isovaline prefers an enantioselective pathway resulting in the production of (*S*)-*sec*-butylamine.

INTRODUCTION

Carbonaceous chondrites are natural reaction vessels carrying within their mineral matrix the chemical fingerprint of the early solar system. Their study provides insights into the prebiotic organic synthesis that occurred in the presolar nebula, the protoplanetary disk, and during asteroidal aqueous and thermal processing.[1,2] Meteoritic aliphatic amino acids have been found in representatives of all carbonaceous chondrite groups, have become the most heavily studied organic compounds in meteorites[3-5] because of their contribution to the prebiotic inventory available for the origins of life on Earth. The organic matter accreted and/or synthesized inside asteroids and comets could have been an important source of substrates for the origin and early evolution of life on Earth (and possibly elsewhere). Amino acids can be chiral (exhibit handedness) and only left-handed amino acids (L-enantiomers) are translated in biology (homochirality). In the absence of a chiral driving force, the abiotic synthesis of amino acids results in racemic mixtures (1:1 ratio of L- and D-enantiomers). The transition from racemic abiotic chemistry to homochiral life remains an unanswered key question in origins-of-life research.[6,7] Some meteoritic amino acids possess L-enantiomeric excess, with the best studied example of these being isovaline (e.g., present at up to 18% L-excess[‡] in the Murchison,[8] and other meteorites)[9]. The presence of non-racemic meteoritic amino acids suggests that homochirality on Earth could have been seeded and developed from the delivery of L-enantioenriched amino acids by carbonaceous chondrites.

Low-temperature aqueous alteration of meteoritic parent bodies may have occurred in melting water ices that resulted from heating caused by radioactive decay of short-lived radionuclides and/or by impacts throughout the history of the asteroid parent body.[10-12] These icy inclusions may have contained the starting materials needed for the formation of meteoritic amino acids through the Strecker synthesis, Michael addition, and/or other synthetic routes.[13-15] Aqueous processes in the parent body of the Murchison and other carbonaceous chondrites belonging to varying petrologic types have been proposed to occur at pH ranging from 6-12,[16] however, this value may greatly depend on the temperature of the system, adsorption of ions, flows of gases and fluids, and the specific mineral composition of the sample, among other parameters.[17] Additionally, continuous heating of these aqueous inclusions may have

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[‡] (*S*)- and (*R*)-notations are more appropriate, however, the amino acid literature uses L- and D-notations based on analogy with glyceraldehyde; for consistency to the nomenclature used in previous meteoritic amino acid studies, we equate (*S*)- to L- and (*R*)- to D- based on their structures.

69   resulted in the decomposition of synthesized amino acids through interactions with liquid water and

70   mineral.[18-20] It has been suggested that the decomposition of amino acids through deamination and

71   decarboxylation driven by aqueous and thermal processes inside the asteroid parent body may result in

72   the formation of carboxylic acids and amines.[21-23] Both aliphatic carboxylic acids and amines have been

73   isolated from carbonaceous chondrites.[21-26] To date, however, the role of amino acids in the production

74   of carboxylic acids and amines through their decomposition under asteroid-like conditions remains poorly

75   understood.[27]

76   In this work, we modeled the effects of asteroidal aqueous processing on the decomposition of

77   amino acids leading to the formation of monocarboxylic acids and monoamines (hereafter called

78   "carboxylic acids" and "amines"). We selected the amino acids glycine and isovaline as proxies for studies

79   of amino acid decomposition under meteoritic conditions using full atomistic molecular dynamics (MD)

80   simulations based on density functional theory (DFT). Glycine is a common achiral amino acid which has

81   been identified in all carbonaceous chondrites and in other extraterrestrial environments, including lunar

82   regolith and comets.[28-31] The chiral amino acid isovaline is rare in the terrestrial biosphere but common in

83   many carbonaceous chondrites and is of interest because excesses of its L-enantiomer have been

84   measured in several meteorites,[8,9,32-34] creating a potential link between its exogenous delivery to Earth

85   and the origins of terrestrial biological homochirality. Our computational study aimed to unveil

86   prospective synthetic relationships between meteoritic amino acids, carboxylic acids, and amines and the

87   maturation of meteoritic organics prior to, during, and after their formation in the asteroid parent body.

88   The key component of this work was assessing the decomposition mechanisms of amino acids in

89   hydrothermal aqueous solution and identifying their decomposition products to establish potential links

90   for the origins of some meteoritic soluble organic compounds. Experimental results on the decomposition

91   of amino acids often produce controversial and, in some cases, contradictory results,[35] leaving unclear

92   whether amino acids would preferentially decompose through decarboxylation or deamination and what

93   the identity of their reaction intermediates are under neutral hydrothermal conditions that may be

94   compatible with those experienced inside the asteroid parent body. Our recently developed approach,

95   which defines a class of general topological reaction coordinates, is particularly well suited for complex

96   multistep mechanisms in aqueous media.[36] We performed *ab initio* molecular dynamics simulations

97   (Born-Oppenheimer) combined with advanced techniques for enhanced sampling of rare events and the

98   reconstruction of free energy landscapes and kinetics (metadynamics, umbrella sampling, weighted

99   histogram analysis, committor analysis)[37-38]. Our simulation clarifies the decomposition products and

100   synthetic origins of prebiotic molecules in meteorites related to glycine and isovaline, shedding new light

101   on fundamental atomic-level mechanisms behind meteoritic prebiotic chemistry.

METHODS

***Ab initio* molecular dynamics and free energy calculations**

We explored the decomposition products of glycine and isovaline using *ab initio* MD, which has well demonstrated its predictive power in describing amino acid synthesis in computer simulations of the Miller experiments.[39] The asteroidal conditions were constrained based on previous meteoritic petrologic analyses, which suggest that aqueous alteration occurred between 0 and 340 °C).[16,40-43] Amino acids may be present as neutral zwitterions inside the parent body based on large abundances of ionic species (cation and anions from mineral complexes and water-soluble salts). To model asteroid-like hydrothermal processing, we selected a temperature of 100 °C and an aqueous environment in which amino acids would be present as zwitterions. We acknowledge that the conditions for hydrothermal processing situates the reactions in bulk water, without explicit consideration of minerals and the presence of species such as a CO_2, CO, and ammonia commonly found in meteorites;[24,44] however, the chosen conditions provide fundamental information on amino acid reaction pathways that may serve as a reference for future experimental and computational works focusing on, for example, the specific effects of particular mineral surfaces on amino acid decomposition.

Born-Oppenheimer MD simulations (at T = 373 K) were performed using interatomic forces computed at DFT level (Perdew-Burke-Ernzerhof functional with Grimme's van-der-Waals corrections).[45,46] Kohn-Sham valence wave functions were expanded on a plane wave basis, as implemented in the CPMD code[47] optimally designed for massively parallel computers. Free energy landscapes were reconstructed employing in tandem metadynamics[48] and umbrella sampling;[49] the former for a fast discovery of reaction pathways and preliminary estimation of barriers, and the latter for a precise sampling of selected pathways. A key component of our approach are path collective variables based on coordination patterns,[38] that allow us to discover unforeseen reaction mechanisms and to obtain reversible pathways without human prejudice about transition states or intermediates. The model consisted of a periodically repeated box including one amino acid molecule (glycine or isovaline) solvated with 79 water molecules. Our simulations followed protocols similar to those previously published,[40] and exploited the plugin Plumed;[50] further details are provided in Supplementary Information. We cannot exclude the existence of amino acid decomposition pathways different from those described in the following sections. However, the reaction mechanisms characterized in this work are the result of extensive *ab initio* simulations, totaling more than 3 ns (80 ps of equilibration trajectories, 300 ps of exploratory metadynamics, 2000 ps of umbrella sampling, and 1000 ps of committor analysis), equivalent

133 to 1.3 million Intel® Xeon® E5-2690v3 hours. For comparison, the majority of published *ab initio* molecular

134 dynamics studies of chemical reactivity are limited to a total simulation length of tens of ps.

RESULTS

Decomposition of amino acids

Glycine and isovaline differ at the alpha-carbon (C_α; Scheme. 1), leading to significant differences in their decomposition pathways; in glycine, the C_α is accessible to the solvent, whereas in isovaline the C_α is protected by a coordination shell of four heavy atoms. We investigated the decomposition of glycine and isovaline through deamination or decarboxylation pathways in simulated asteroidal hydrothermal processing under aqueous neutral conditions, the investigated mechanisms are shown in Figure 1 panels a and b. Additionally, we show the schematic summaries of the step-by-step reaction mechanisms for glycine and isovaline decompositions in Figures S1, S2, and Movies S1-S6 available in the Supplementary Information. Corresponding free energy profiles and landscapes are shown in Figure 2 and Figure 3 respectively, and representative transition state and electronic structures are depicted in Figures 4 and 5.

The two decomposition mechanisms of solvated glycine (**1**) and isovaline (**5**) start from their favored zwitterionic species (Figure 1). The deamination of glycine does not show the formation of acetate (**2'**), which is the simplest aliphatic deamination product one could envisage; indeed, the loss of neutral ammonia from the glycine zwitterion would leave a carbocation species that, formally, would need the addition of a hydride (H^-) to form acetate. Note, however, that this reaction is not viable in practice: even if a hydride were present in solution, it would probably undergo a different fate (reduction of the glycine carboxylate group; reaction not shown here). Instead, the deamination of glycine results in two possible end products: neutral glycolic acid (**2**) and ammonia, or their conjugate base/acid glycolate (**3**) ammonium salt (Figure 1A). The free energy landscape reveals an identical barrier of 40 kcal/mol for these two deamination results (Figures 2a, 3a, 3b, and Table S1 which shows a summary of the simulated reactions free energy barriers in kcal/mol). However, the two possible end states have different stabilities, with the charged species being lower in free energy by 5 kcal/mol, as expected from the pK_a of glycolic acid (3.83 at 300 K) and from the pK_b of ammonia (4.7). The configuration of the transition state was identified by committor analysis, and a representative structure is shown in Figure 4a. The transition state is characterized by the partial breaking of the C–N bond with simultaneous splitting of a water molecule and hydroxide (OH^-) nucleophilic attack at the C_α (with partial carbocation character; Figures 4a and 5a).

The free energy landscape for glycine decarboxylation is depicted in Figure 3c. The transition state (Figures 4b and 5b) is characterized by the simultaneous i) partial breaking of the C–C bond, ii) partial linearization of the CO_2 group, and iii) a half-split water molecule, whose proton moves towards the C_α atom (with partial carbanion character). We estimated a free energy barrier of 43 kcal/mol, very similar

166 to that found for deamination. The product state appears 12 kcal/mol less stable than reactants; however,

167 we note that the further reaction $[CO_2 + OH^- \rightarrow HCO_3^-]$ is expected to decrease the free energy by 10

168 kcal/mol,[51] bringing glycine to a similar stability with respect to $[CH_3NH_3^+ + HCO_3^-]$ (Figure 2a, Table S1).

169   Compared to glycine, isovaline (**5**) displays a similar decarboxylation pathway, leading to the

170 formation of *sec*-butylamine (**9**), but a different deamination pathway resulting in products (**6** and **7**)

171 which in turn may decarboxylate to yield 2-butene (**8**; Figures 1b and 2b). Similar to the deamination of

172 glycine, no presence of the unsaturated deaminated analog, 2-methyl butanoate (**7′**) was observed as a

173 deamination product of isovaline (the same argument observed above for glycine can be invoked here).

174 The deamination of isovaline yields two products as the possible outcomes of the same transition state,

175 corresponding to a barrier of 36 kcal/mol (Figures 2b, 3d, 4c, 5c). The first product is 2,2-ethyl-methyl-α-

176 acetolactone (**6**) (albeit with a C–O–C bridge that is intermittently formed and broken on the picosecond

177 time scale), structurally simple to reach from the transition state by reducing the C–C–O angle. This species

178 is, however, quite unstable, with a free energy 32 kcal/mol higher than isovaline (Figures 2b, 3d). A more

179 stable product, lying at 20 kcal/mol above isovaline, is *cis*-2-methyl-2-butenoic acid (commonly known as

180 angelic acid, **7**) in its carboxylate anionic form, obtained after deprotonation (operated by the leaving

181 ammonia molecule) of the methylene group with concomitant formation of a C=C double bond. We

182 remark that a single transition state (labeled $TS_{5(6,7)}$ in Figures 2-5) leads to compounds **5**, **6**, **7** with equal

183 probability, as we explicitly tested with committor analysis. Note that the *trans* isomer of **7**, known as

184 tiglic acid, could also be obtained by a similar mechanism, however we did not explicitly perform the latter

185 simulation. Experimentally, pure angelic acid tends to irreversibly convert to tiglic acid, which is more

186 stable by 8.5 kcal/mol than angelic acid as deduced by heats of combustion.[52] Therefore, at 100 °C, our

187 simulations suggest that an isovaline decomposition to angelic/tiglic acid will result in an equilibrium ratio

188 of about 10^{-7} between products and reactants.

189   We extended our metadynamics simulations to explore possible further evolution of angelic acid

190 (**7**) toward decarboxylation, by explicitly targeting this reaction with extensive umbrella sampling

191 simulations. Starting from **7**, a barrier of 55 kcal/mol leads to the cleavage of the C–COO⁻ bond and

192 protonation of the C_α from a water molecule, eventually resulting in the formation of carbon dioxide and

193 2-butene (Figures 2b and 3e,). The high barrier may be explained by the relative inaccessibility of the C_α

194 to solvent molecules, due to surrounding bulky groups, until a significant elongation of the C–COO⁻ bond

195 is achieved (Figures 4d and 5d). We remark that if, instead, decarboxylation is attempted starting from

196 2,2-ethyl-methyl-α-acetolactone (**6**), the system rapidly evolves first into angelic acid, confirming that this

197   lactone is an unlikely deviation, apparently of scarce relevance in the reconstructed reaction network. Our

198   results suggest that decarboxylation of angelic acid entails a very large free energy barrier, and results in

199   products of very low stability compared to isovaline; thus, it is not expected that 2-butene derived from

200   isovaline would significantly contribute to the observed pool of meteoritic organic molecules.

201   Direct isovaline decarboxylation proceeds through a simple mechanism (Figure 1), yielding *sec*-

202   butylamine after crossing a barrier of 48 kcal/mol (Figures 2b, Table S1). This decomposition channel is

203   potentially relevant: compared to deamination, decarboxylation encompasses a much higher barrier (with

204   half-life of 10^7 years at 100 °C or 10^4 years at 150 °C, as estimated from Eyring's transition state theory

205   formula) but, in turn, it leads to a much more stable product (only 2 kcal/mol above isovaline; Figures 2b

206   and 3f). Similar to the reaction connecting angelic acid (**7**) with 2-butene, the mechanism involves the

207   cleavage of the C–COO$^-$ bond and the protonation of the C_α atom from a water molecule (Figures 4e and

208   5e). At the transition state, the C_α keeps a tetrahedral sp^3 character, enabling access to solvent only on

209   the same side as the leaving CO_2 moiety and protecting the opposite side by steric hindrance. This is

210   confirmed by inspection of all the available trajectories in the transition state region, including umbrella

211   sampling and committor analysis. Thus, the solvent proton replaces the carboxylate functional group

212   without changing the chirality of the C_α. We conclude that the decarboxylation pathway itself does not

213   lead to *sec*-butylamine racemization. We note, instead, that the presence of the α-hydrogen on *sec*-

214   butylamine is generally accepted to strongly increase its hydrothermal racemization rate. In our specific

215   case, this indicates that *sec*-butylamine derived from isovaline would racemize at a faster rate than its

216   parent amino acid inside the asteroid parent body. Preliminary metadynamics simulations suggested

217   other possible decarboxylation mechanisms, such as simultaneous isovaline decarboxylation and

218   deamination, or the formation of formic acid instead of carbon dioxide; however, such pathways turned

219   out to be unlikely under the scrutiny of umbrella sampling.

DISCUSSION

Comparison to previous computational modeling and experimental results

220

221

222   The current understanding of amino acid decomposition processes in solution is mostly based on

223   experimental information. The most relevant theoretical study to date addressed glycine and alanine

224   decomposition based on a quantum mechanics / molecular mechanics (QM/MM) approach, with the

225   amino acid and up to two water molecules described with a semi-empirical molecular orbital method ,

226   immersed in a bath of classical water molecules.[53] The previous work yielded a glycine decarboxylation

227   barrier of 46 kcal/mol, close to ours. In our simulations, based on a more accurate quantum description

228   and on more general collective variables, we find that, with respect to the previous work, the products

229   are significantly less stable than the reactants. Concerning glycine deamination, Alexandrova and

230   Jorgensen (2011) reported a free energy barrier of 51 kcal/mol for the formation of an α-lactone, and >

231   90 kcal/mol for the formation of glycolic acid. Instead, we predict glycolic acid formation as the only viable

232   reaction, with a much lower barrier of only 40 kcal/mol, in good agreement with experiments as discussed

233   below, thus pointing to a more accurate description of the reaction pathway in our framework.

234   The interpretation of the available experimental results in the literature is controversial. As

235   pointed out by Cox and Seward (2007), various reports using different experimental conditions and

236   reaction vessels often presented contradictory conclusions, so that it is not clear whether glycine and

237   other amino acids preferentially decarboxylate or deaminate, and through which mechanism. In several

238   cases, decarboxylation was *a priori* assumed to happen, and deamination was not explicitly tested. A

239   decarboxylation enthalpy barrier of 39 kcal/mol above 170 °C was inferred from NMR analyses,[54] and

240   similarly from chromatography above 200 °C.[51] Snider and Wolfenden (2000) also reported a slow-down

241   of the reaction with increasing ionic strength, suggestive of a transition state less polarized than the

242   reactants, which is consistent with the decomposition mechanisms found in this work. An enthalpy barrier

243   of 33 kcal/mol was inferred from an *in situ* infrared spectrometer flow reactor above 310 °C.[56] The free

244   energy barriers we observed for deamination and decarboxylation are similar within statistical errors

245   (about 2 kcal/mol), and using Eyring formula ($k = k_BT/h \ e^{-\Delta G^*/k_BT}$) the corresponding decomposition rate is

246   between 10^{-11} and $10^{-13}\,s^{-1}$ at 100 °C, equivalent to half-lives between one thousand and several tens of

247   thousands of years; the upper rate is in agreement with experimental measurements.[54-56] Given the

248   survival of amino acids in meteorites to the present day, these half-lives would be consistent with

249   suggestions that although there may have been available liquid water for several millions of years inside

250   the parent body, the alteration processes may have only been episodic and not continuous throughout

251   that period of time (i.e. episodes of aqueous alteration may have been triggered by short-term events like

252   impacts).[42]

253   Similar to chromatographic experiments,[55] and based on our findings, we conclude that

254   deamination and decarboxylation decomposition pathways could be simultaneously available in a

255   hydrothermal glycine solution. The simulations, however, predict the formation of ammonium glycolate

256   and methylammonium bicarbonate ions at a similar pace, with the latter products being much more stable

257   than the former (if the exothermic hydroxylation of CO_2 occurs starting from methylammonium) and in

258   similar equilibrium proportions as the glycine reactants. From the kinetic point of view, deamination

259   appears to be moderately favored over decarboxylation, even if the free energy difference of 3 kcal/mol

260   does not appear very significant considering statistical uncertainties as well as the intrinsic accuracy of

261   DFT calculations.[57] Therefore, following long-term aqueous conditions on asteroid parent bodies,

262   methylamine is expected as the dominant meteoritic glycine decomposition product (leaving aside the

263   possibility of products themselves entering further reaction pathways). Indeed, the abundance of amines

264   is greater than that of amino acids in meteorites experiencing extended increasing aqueous processing,

265   while the opposite is true in minimally aqueously altered carbonaceous chondrites.[23-26] It should be noted

266   that *in situ* spectrophotometric measurements pointed to sizable glycine dimerization in addition to

267   decomposition.[35] However, investigating such additional multi-molecular process is beyond the scope of

268   this work.

269

270   **Constraints of our asteroid-like processing model**

271   Our work provides detailed quantitative insight about amino acid decomposition reactions in

272   simulated asteroidal hydrothermal conditions; however, discussions of the interpretation and relevance

273   of these results from the point of view of asteroid parent body processes and meteorite organics must

274   consider the following factors. First, we chose a specific temperature and two amino acids particularly

275   representative of meteoritic chemistry: the simplest and most common one (glycine), and the one with

276   the most reports of L-enantiomeric excess in meteorites (isovaline). Temperatures above 100 °C and a

277   higher concentration of protons or catalytic transition metal ions, which may have occurred for

278   extensively altered chondrites, may strongly alter the amino acid decomposition kinetics. Second, the

279   reaction network we reconstructed (Figure 1) is part of a larger scheme in which amino acid

280   decomposition products can take part in further reactions with other species not explicitly considered

281   here. In this sense, even a relatively evanescent intermediate, of apparent low relevance, could increase

282   the available molecular inventory, and possibly play a key role toward a stable product in the more global

283   prebiotic network in the parent body. Third, the reaction rates and equilibrium populations we inferred

284   are pertinent to bulk, neutral water solutions. As such, they provide a reference reaction network that

285   can be used as a starting point to address various questions about the viability of different prebiotic

286   scenarios, including: what is the optimal range of temperature for a given reaction?; would an acidic or

287   alkaline pH favor a synthetic/decomposing pathway?; under what conditions would mineral surfaces play

288   significant role in stabilizing transition states, intermediates, and products? These are questions that

289   future experimental and theoretical efforts may be able to answer, incrementing our understanding about

290   the origins of prebiotic meteoritic organic matter.

291

292   **Implications for the origins of some meteoritic soluble organic compounds**

293   Although meteoritic organic compounds formed in distinct stages through the birth and

294   evolution of the solar system, we can constrain their origin to two main environments and regimes: (1)

295   pre-parent body, dominated by gas- and ice-grain chemistry that occurred in the molecular cloud, the

296   solar nebula, or the protoplanetary disk, and (2) parent body, dominated by hydrothermal processes.

297   Given the unknown concentration of the molecular species available for chemical development, the

298   unsettled and changing physical conditions such as temperature and pressure, and the uncertain level of

299   processing that occurred in each parent body, it may prove challenging to assess the dominant synthetic

300   routes leading to the origins of meteoritic amino acids.[27] However, it is useful to note that (a) the detection

301   of the simplest amino acid (glycine) in the interstellar medium remains controversial,[58,59] (b) the observed

302   abundances of meteoritic amino acids and those synthesized after hydrolysis of UV-irradiated interstellar

303   ice analogs decreases with increasing carbon,[60,61] and (c) only aminoacetonitrile, but no glycine, has been

304   observed before the hydrolysis of interstellar ice analogs irradiated at temperatures ranging from 20 to

305   300 K.[62] From these observations, it can be suggested that some degree of processing inside the parent

306   body may favor augmenting the molecular abundance and diversity of meteoritic amino acids, and that

307   prolonged processing could result in the amino acid decomposition pathways studied here.

308   From our computational analyses, we have found that the deamination and decarboxylation of

309   glycine occurs with similar barriers leading to the production of the hydroxylated acid (glycolic acid) and

310   methylamine. In contrast, the deamination and decarboxylation of isovaline occurs at different rates, and

311   it is not the hydroxylated acid, but rather the α-β-unsaturated acid (**7**) that would be the fastest

312   decomposition product over *sec*-butylamine, although the amine is expected to dominate at equilibrium

313   by virtue of its higher stability. The difference in decomposition products may originate from the relative

314   inaccessibility of the C_α to solvent molecules in isovaline relative to glycine (see Scheme 1 and electron

315   localization function – ELF - analysis shown in Figure 5). It should be noted, furthermore, that neither the

316   decarboxylation of glycine nor that of isovaline result in the formation of formic acid. This observation

317   suggests that meteoritic formic acid must have formed before the accretion of the parent body, or through

318   reaction mechanisms that are not related to the decomposition of amino acids inside the parent body,

319   e.g., catalytic hydrogenation of CO_2, or the hydrolysis of the methyl formate produced by the reaction of

320   methanol and carbon monoxide.

321   Our results do not provide evidence for the formation of saturated aliphatic monocarboxylic

322   acids, namely, acetic acid and 2-methylbutanoic acid, as deamination products of glycine and isovaline

323   respectively. These carboxylic acids have been detected in various carbonaceous chondrites and

324   constitute some of the most abundant water-soluble organic compounds in meteorites,[22,24,25,63] but our

325   calculations suggest that deamination of amino acids does not contribute to their reported abundances.

326   This contrasts with methylamine and *sec*-butylamine, the corresponding decarboxylation products of

327   glycine and isovaline. This observation suggests that either aliphatic acids originated in cold interstellar

328   pre-solar regions and were later incorporated inside the parent body, or that they formed from organic

329   species that underwent synthetic processes other than glycine and isovaline decomposition, e.g.,

330   oxidation of aliphatic aldehydes or hydrolysis of alkyl nitriles inside the parent body.

331   Our calculations show an α-β-unsaturated compound (angelic acid, **7**) as the deamination product

332   of isovaline, with a further decarboxylation leading to 2-butene. The abundance of **7** (or any other α-β-

333   unsaturated monocarboxylic acid) has not been evaluated in analyses of meteoritic organics. Although 2-

334   butene has been reported in analyses of extracts from the Murchison meteorite in a mix of isomers,[24] this

335   decomposition pathway exhibits a very high barrier that suggests that the meteoritic origins of 2-butene

336   would not be necessarily related to the decomposition of the amino acid, except in cases of extended

337   hydrothermal processing at higher temperatures. Another possibility is that **7** could react to form β-amino

338   acids inside the parent body, considering that α-β-unsaturated carboxylic species readily undergo 1,4-

339   nucleophilic additions.[64-65] Indeed, the concentration of β-amino acids is higher than that of α-amino acids

340   in more extensively aqueously altered carbonaceous chondrites.[34,66] Our computational results of the

341   deamination of isovaline (an α-alkyl-α-amino acid) would support a parent body scenario in which the

342   concentration of β-amino acids increases at the expense of the decomposition of α-amino acids through

343   aqueous alteration. However, the possibility that the higher abundance of β-amino acids may be an
344   inherent characteristic of the parent body's molecular budget might not be discarded.

345   The more stable decarboxylation product of isovaline, *sec*-butylamine, has been found in
346   carbonaceous chondrites exhibiting varying levels of aqueous and thermal processing.[28,67] The stability of
347   *sec*-butylamine is consistent with the idea that at least a portion of this meteoritic amine may be produced
348   upon decomposition of the amino acid either in the parent body stage or through amino acid extraction
349   (the extraction of amino acids from meteoritic samples is typically done using hot water at 100 °C) and
350   sample processing during analyses.

351   In addition to the detailed decomposition pathways shown here and reading the reactions in
352   Figure 1 in reverse order, it can be predicted that glycine could be obtained from methylamine or from
353   glycolic acid, and isovaline from *sec*-butylamine or from angelic acid, through relatively-low free energy
354   barriers that range from 15 to 45 kcal/mol. These results suggest that the synthesis of meteoritic amino
355   acids under hydrothermal conditions appear feasible even without invoking the aid of catalyzers or
356   synthetic routes such as the Strecker-cyanohydrin synthesis, similar to recent experimental results.[15] Our
357   predictions could be tested in future experiments, helping to inform both terrestrial and extraterrestrial
358   amino acid formation processes.

359

360   **Enantiomeric excess at the dawn of the solar system**

361   The enantiomeric excess (*ee*) of L-isovaline has been observed in meteorites with various
362   processing histories,[8,9,33,34] and a direct relationship has been suggested between the degree of aqueous
363   alteration and the percentage of L-*ee* found in carbonaceous chondrites.[8] The enantiomeric composition
364   of meteoritic *sec*-butylamine has shown contradictory results. Aponte et al. (2015, 2016), reported
365   racemic mixtures of *sec*-butylamine in carbonaceous chondrites from which enantiomeric excesses of L-
366   isovaline was found ranging from 0 to 18%. In contrast, enantiomeric enrichments of (*S*)-*sec*-butylamine
367   ranging from 0 to 66% in moderately and minimally aqueously altered CR2 and CR3 chondrites in which
368   isovaline was racemic (L-*ee* from 0 to 3 ± 4%[34,68]) was also reported.[69] One notable piece of information
369   obtained from our models is the enantioselective decomposition of isovaline into *sec*-butylamine.
370   According to this result, any *ee* present in L-isovaline, would be conserved when decarboxylated into (*S*)-
371   *sec*-butylamine, indicating that an observed (*S*)-*ee* of *sec*-butylamine could be a potential indicator of L-
372   isovaline *ee*. Additionally, our results suggest that *sec*-butylamine will racemize at a faster rate than

373   isovaline given the presence of the α-hydrogen and the lower steric hindrance; in fact, isovaline is more

374   stable against racemization relative to α-H-α-amino acids.[70]

375   The lack of observed *ee* of *sec*-butylamine in meteorites across varying degrees of alteration and

376   in meteorites with a significant L-isovaline *ee* may suggest that, despite our theoretical results, the

377   decarboxylation of enantioenriched isovaline is not a major reaction in the parent body. However,

378   alternate explanations exist, including that: (a) enantioenriched *sec*-butylamine is not observed because

379   of the expected faster racemization of the amine relative to that of isovaline, or (b) a small L-*ee* of isovaline

380   existed that created an *ee* in *sec*-butylamine formed upon decomposition that is below current detection

381   limits, but that parent-body conditions subsequently amplified the isovaline *ee*,[71,72] without having a

382   significant effect on the enantiomeric composition of *sec*-butylamine. The (*S*)-*sec*-butylamine *ee* reported

383   by Pizzarello and Yarnes (2016) would suggest that *sec*-butylamine and isovaline may be synthetically

384   linked through the decomposition of the amino acid in those less-aqueously-altered carbonaceous

385   chondrites.

386 **CONCLUSIONS**

387   In this work, we modeled the effects of aqueous processing leading to the decomposition of the
388 amino acids glycine and isovaline through deamination and decarboxylation. We have been mindful of
389 not being contradicting or confusing towards the broader scientific community in our mechanistic
390 assessments. Our work aimed to be a stepping point for future work that may link computational and
391 experimental analyses to the origins of organic matter in the solar system; no changes were made to the
392 text. Our results have several implications for the origins of soluble meteoritic organic compounds such
393 as amino acids, carboxylic acids and amines. We found that the deamination and decarboxylation of
394 glycine occur through similar barriers, suggesting that both processes would occur in parallel inside the
395 parent body. In contrast, we found that the deamination of isovaline occurs at a faster rate than
396 decarboxylation, and that this process results in the formation of an α-β-unsaturated acid. This may help
397 to explain the observed increase in relative abundance of β-amino acids compared to α-amino acids in
398 meteorites with increasing levels of aqueous processing. Additionally, although aliphatic amines were
399 found as decarboxylation products of glycine and isovaline, our computations did not find evidence for
400 the formation of analogous aliphatic carboxylic acids after deamination, suggesting that these compounds
401 may have synthetic origins that are not related to the decomposition of amino acids inside the parent
402 body. Finally, we found that the decarboxylation of isovaline occurs in an enantioselective fashion,
403 meaning that the chiral configuration of isovaline would be conserved in the *sec*-butylamine produced.
404 The decomposition of isovaline in meteorites exhibiting L-*ee* of this amino acid would produce (*S*)-
405 enantioenriched *sec*-butylamine, although the enantioenriched amine could then racemize more quickly
406 than its amino acid precursor.

Acknowledgments

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417

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592   **Scheme 1.** Structures of glycine and isovaline zwitterions (color highlights the amine [blue, NH_2] and acid

593   [red, CO_2H] functional groups). Illustration of the different degree of steric hindrance around the alpha-

594   carbon (C_α).

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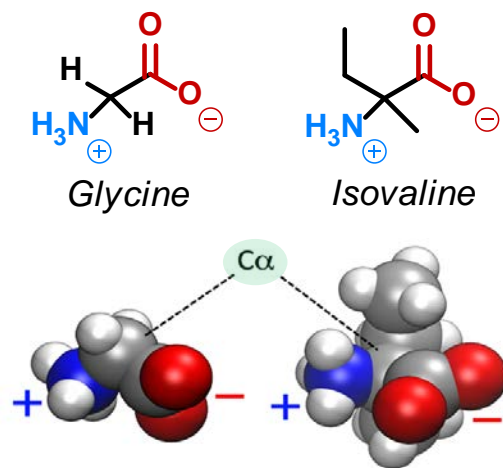
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598 **Figure 1.** Decomposition reactions of glycine (A) and isovaline (B) evaluated in this work. All simulations
599 are performed in bulk water solution at T = 373 K. Reactions marked with an X never occur in our
600 simulations: see the related discussion in the Results section. We equate (*S*)- to L-enantiomers for
601 consistency to the nomenclature used in previous meteoritic amino acid studies.
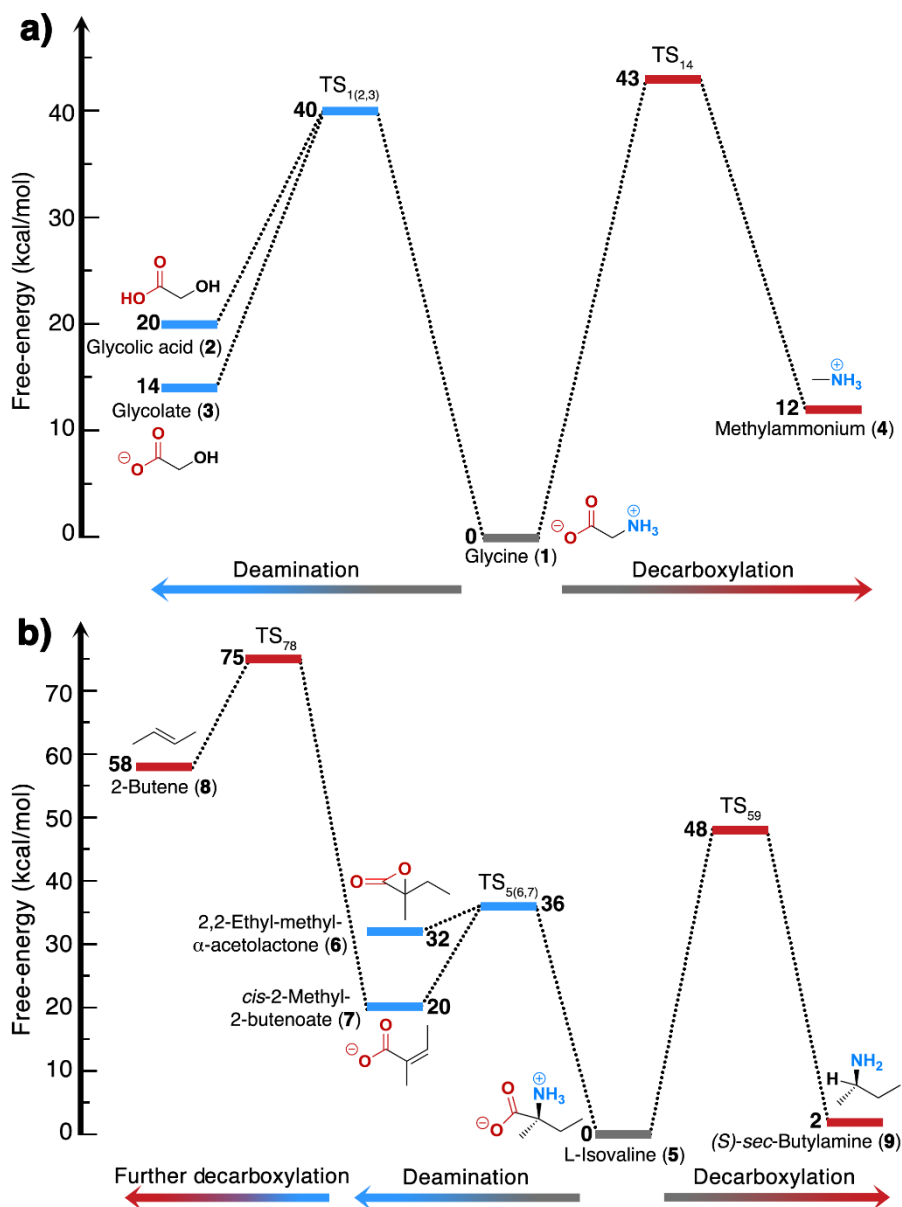
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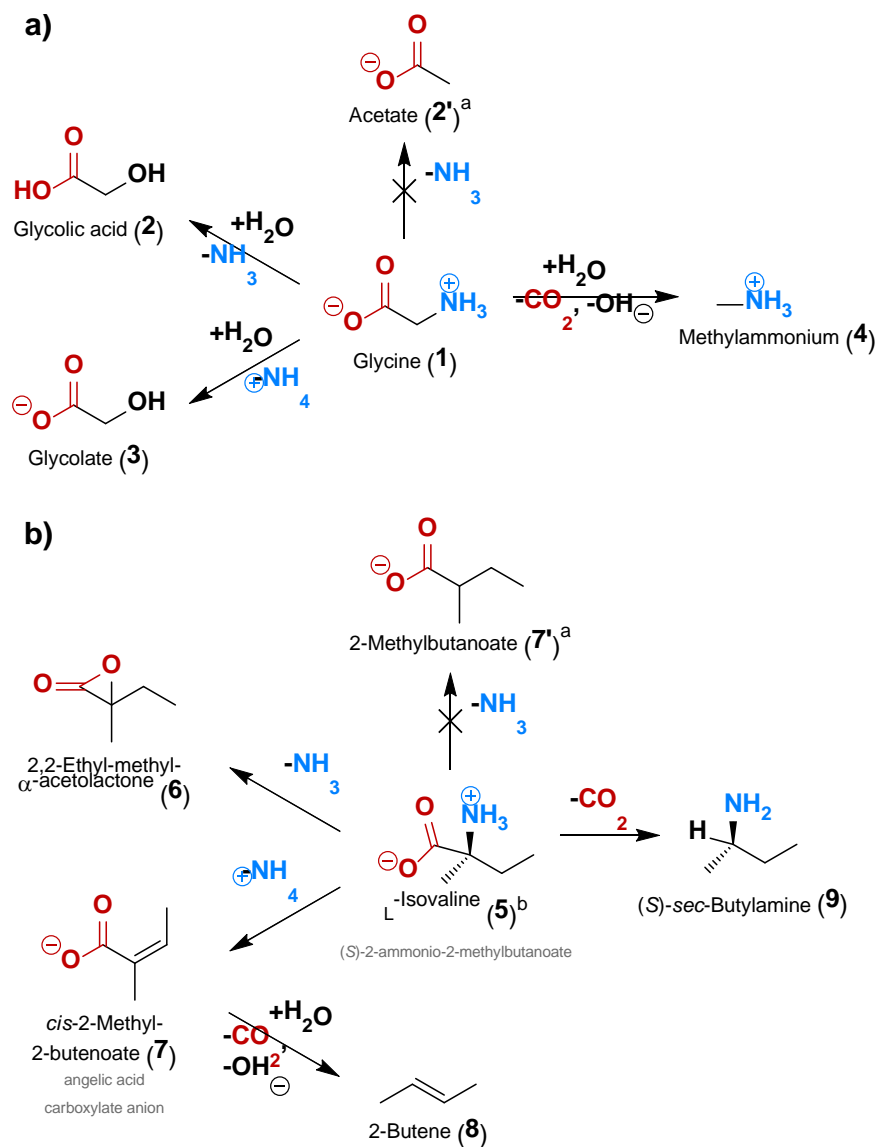
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605 **Figure 2.** Schematic free-energy profiles of a) glycine and b) isovaline hydrothermal decomposition
606 processes from *ab initio* simulations. Statistical uncertainties are estimated as ±1 kcal/mol.

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Figure 3. Free-energy landscapes of chemical reactions at *T*=373 K as a function of path collective

variables based on distances between atomic coordination patterns[36]. The landscapes are obtained from

weighted histogram analysis applied to *ab initio* umbrella sampling simulations, a) glycine deamination

to (**2**), b) glycine deamination to (**3**), c) glycine decarboxylation to (**4**), d) isovaline deamination to (**7**), e)

(**7**) decarboxylation to (**8**), and f) isovaline decarboxylation to (**9**).
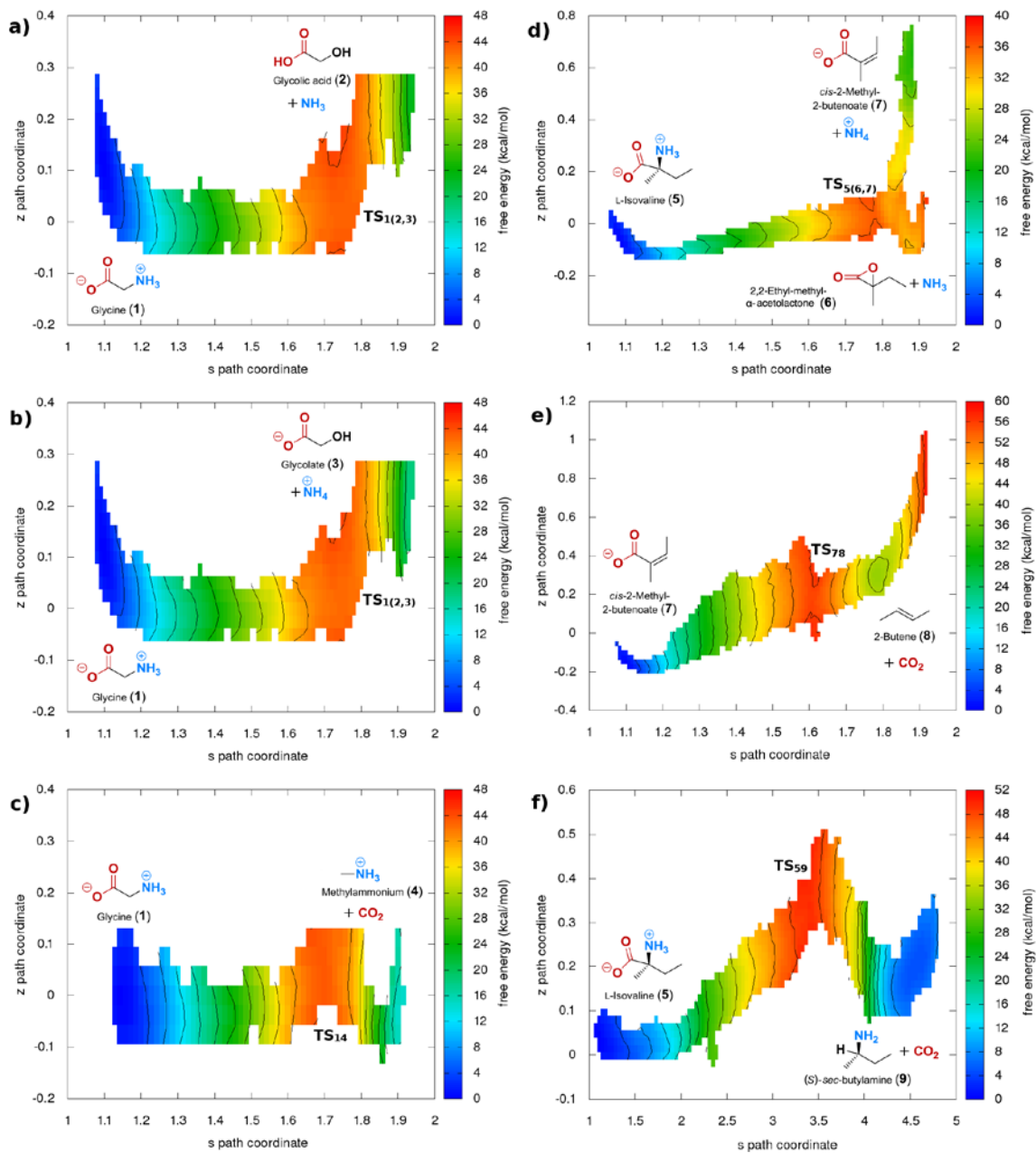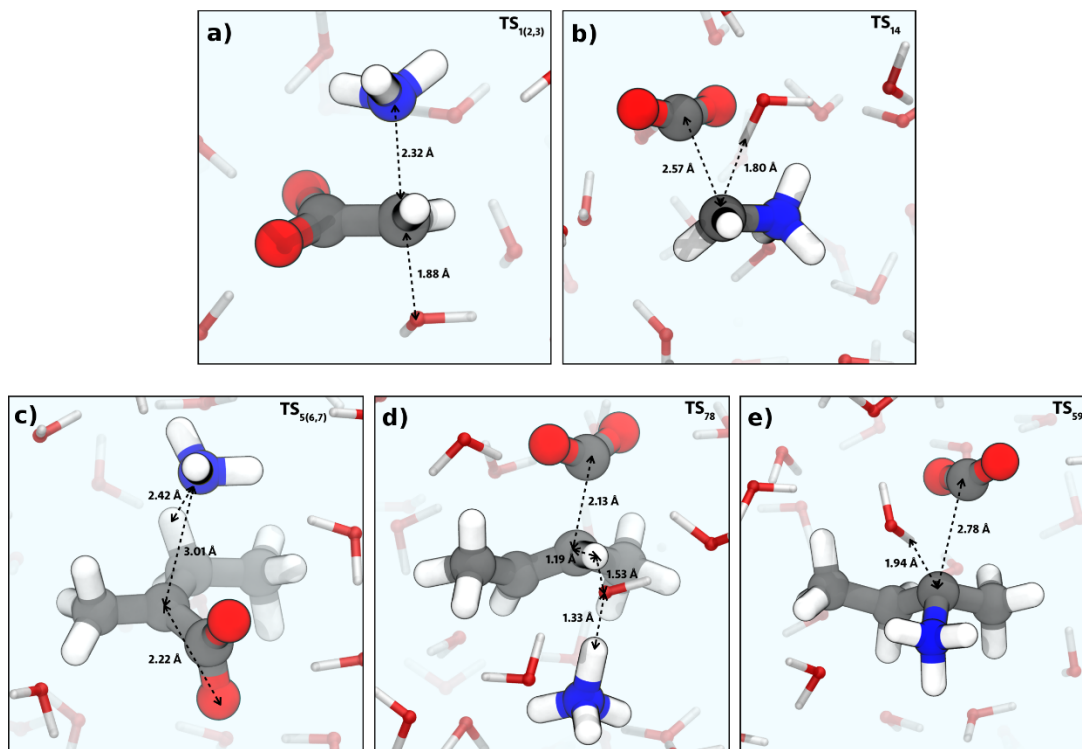
617 **Figure 4.** Representative transition state structures, identified by means of committor analysis on

618 umbrella sampling trajectories for a) glycine deamination, b) glycine decarboxylation, c) isovaline

619 deamination, d) angelic acid decarboxylation, and e) isovaline decarboxylation.
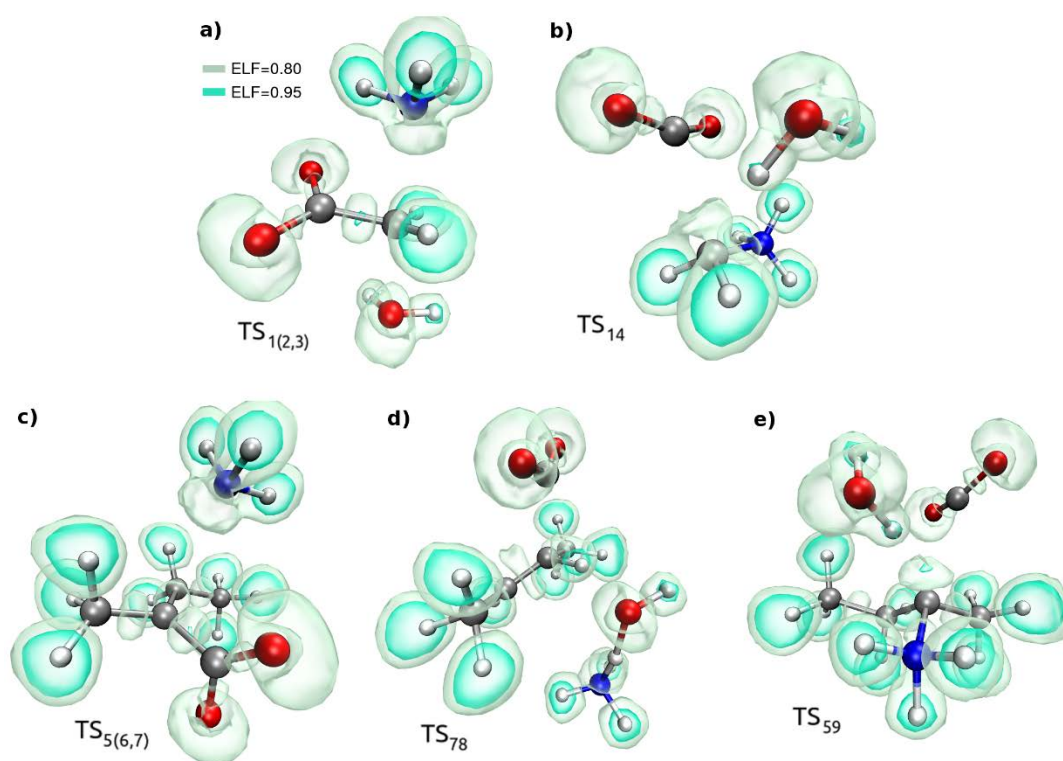
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622   **Figure 5.** Electronic structure of the transition state configurations, represented by means of different

623   isovalues of the electron localization function.[73] Solvent water molecules are present in the calculations

624   but not depicted here, except for those undergoing bond breaking/formation. The corresponding

625   reactions are a) glycine deamination, b) glycine decarboxylation, c) isovaline deamination, d) angelic acid

626   decarboxylation, and e) isovaline decarboxylation. The carbocation character of the C_α is evident in panels

627   a) and c), whereas its carbanion character is evident in panels b) and e).

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629