Nature of the Band Gap and Origin of the Conductivity of PbO$_2$ Revealed by Theory and Experiment

David O. Scanlon, Aoife B. Kehoe, and Graeme W. Watson
School of Chemistry and CRANN, Trinity College Dublin, Dublin 2, Ireland

Martin O. Jones and William I. F. David
ISIS Facility, Rutherford Appleton Laboratory, Chilton, Didcot, Oxon OX11 0QX,
UK Department of Chemistry, Chemistry Research Laboratory, University of Oxford, Mansfield Road, Oxford OX1 3TA, United Kingdom

David J. Payne, Russell G. Egdell, and Peter P. Edwards
Department of Chemistry, Chemistry Research Laboratory, University of Oxford, Mansfield Road, Oxford OX1 3TA, United Kingdom

Aron Walsh
Centre for Sustainable Chemical Technologies and Department of Chemistry, University of Bath, Claverton Down, Bath BA2 7AY, United Kingdom

(Received 6 October 2011; published 7 December 2011)

Lead dioxide has been used for over a century in the lead-acid battery. Many fundamental questions concerning PbO$_2$ remain unanswered, principally: (i) is the bulk material a metal or a semiconductor, and (ii) what is the source of the high levels of conductivity? We calculate the electronic structure and defect physics of PbO$_2$, using a hybrid density functional, and show that it is an $n$-type semiconductor with a small indirect band gap of $\sim$0.2 eV. The origin of electron carriers in the undoped material is found to be oxygen vacancies, which forms a donor state resonant in the conduction band. A dipole-forbidden band gap combined with a large carrier induced Moss-Burstein shift results in a large effective optical band gap. The model is supported by neutron diffraction, which reveals that the oxygen sublattice is only 98.4% occupied, thus confirming oxygen substoichiometry as the electron source.

DOI: 10.1103/PhysRevLett.107.246402 PACS numbers: 71.20.–b, 61.72.S-, 82.47.Cb

Since the invention of the lead-acid battery in 1860, [1] lead dioxide has been of great practical importance due to its utility in many electrochemical processes. [2] These applications depend strongly on the electrical conductivity and redox chemistry of the material [2,3]. While high levels of conductivity are well established, the electronic structure of PbO$_2$ has been a source of controversy in the literature, and to date remains uncertain. It has been postulated that PbO$_2$ is intrinsically metallic [4–9], but conversely, the carrier concentrations have been reported to vary strongly with oxygen partial pressure [10], indicating a mechanism of lattice defect controlled conductivity. Whereas previous band structure calculations based on density functional theory (DFT) found that the conduction band of PbO$_2$ overlaps with the top of the valence band (semimetallic) [7,11], optical band gaps have been reported in the literature of the order of 1.4–2.9 eV [12,13]. It is clear that there is a pressing need for a careful reexamination of the physical properties of PbO$_2$.

In addition to the electronic structure, the role of lattice defects in PbO$_2$ has also been a cause for much debate [14,15]. PbO$_2$ typically displays degenerate $n$-type conductivity, with undoped samples possessing carrier concentrations of $\sim$10$^{19}$–10$^{21}$ cm$^{-3}$ [7,10], depending on sample preparation conditions. This conductivity has been attributed to oxygen vacancies ($V_O$) [16], lead interstitials (Pb$_i$) [17], and adventitious incorporation of hydrogen into the system [12,15,18].

In the lead-acid battery, it has been reported that PbO$_2$ electrodes are substoichiometric on the $cation$ sublattice, and not on the anion sublattice, based on neutron diffraction experiments [14,15,19]. Even recent studies refer to the fact that the oxygen sublattice is “complete” [2]. The presence of lead vacancies ($V_{Pb}$) as the dominant defect species should cause a considerable number of hole carriers in the system, although $p$-type PbO$_2$ has never been reported. In isostructural and isoelectronic SnO$_2$, the formation energy of a cation vacancy is much higher than that of the anion vacancy [20,21], and a similar behavior would be expected here.

In this Letter, through a combination of first-principles band structure calculations and experimental characterization of PbO$_2$, we present the first consistent explanation for the material behavior. PbO$_2$ has a small but finite band gap and a light electron effective mass. The calculated charge neutrality level lies deep in the conduction band, which is consistent with high levels of electrical conductivity. Conduction electrons are created through oxygen loss,
which is an exothermic process in the noninteracting limit, as confirmed by neutron diffraction measurements.

Computational Method.—All DFT calculations were performed using the VASP code [22]. Interactions between the core and valence electrons described within the scalar-relativistic projector augmented-wave (method) [23]. The calculations were performed using the Perdew-Burke-Ernzerhof (PBE) [24] exchange-correlation functional augmented with 25% screened Hartree-Fock exchange, producing the Heyd-Scuseria-Ernzerhof (HSE06) functional [25]. HSE06 has been successful in reproducing the structural and band gap data for a wide range of metal oxide systems [26–35]. Plane wave cutoff and k-point sampling were tested, with a cutoff of 400 eV and a k-point sampling of 6 × 6 × 4 for the 6 atom unit cell of PbO2 found to be sufficient. The ionic forces were converged to less than 0.01 eV Å−1. The optical transition matrix elements, calculated following Fermi’s golden rule, were used to construct the imaginary dielectric function and the corresponding optical absorption spectrum [36]. Point defects were calculated in their various charge states, as outlined by Freysoldt et al. [39], using an isotropic dielectric constant ε0 = 10. Band filling by resonant defect levels was treated using the standard approach [40]. The chemical potential limit of the O-poor conditions was checked, with PbO found to be a competitive phase.

Experimental Method.—Powder neutron diffraction was performed on a 5 g sample of commercially available PbO2 powder (Sigma Aldrich, U.K., 99.999% pure) at the ISIS Spallation Neutron Source facility, Rutherford Appleton Laboratory, U.K. Samples were held in vanadium cans and all data collected on the Polaris diffractometer [41]. The diffraction data were fitted by the Rietveld method using the TOPAS Academic program suite [42], and is provided as Supplemental Material [43]. Two phases of PbO2 were identified, a minority α-phase (10.4 wt%, space group Pbcn, a = 5.001 Å, b = 5.931 Å, and c = 5.443 Å) and a majority β-phase (89.6 wt%, space group P42/mnm, a = 4.951 Å, and c = 3.383 Å). For β-PbO2 the oxygen occupancy and oxygen x and y site parameters were refined independently, as were the unit cell lattice parameters and the anisotropic thermal parameters for Pb and O. The peak profiles were modeled with Lorentzian and Gaussian functions and a weighted-profile reliability factor of 1.88% achieved. Analysis of the atomic occupancies confirmed an oxygen deficiency of approximately 1.6%, with an oxygen occupancy of 0.984.

Geometry and Electronic Structure.—PbO2 (mineral name plattnerite) crystallizes in the rutile structure (β-phase), in which Pb is coordinated to six oxygen in a distorted octahedron (4 short bonds, 2 long bonds, D4h symmetry) and each oxygen is coordinated to three Pb, via two long bonds and one short bond. The calculated structural and band gap data for PbO2 are shown in Table I. For comparison we calculated the band structure using both the hybrid HSE06 functional and the semilocal PBE functional. The PBE method overestimates the experimental lattice constants by ~2%, whereas the HSE06 method results in a structure in excellent agreement. The major difference between the PBE and HSE06 electronic structure is that the HSE06 functional predicts PbO2 to be a semiconductor with a finite band gap, and the PBE predicts a metallic material with an overlap of 0.7 eV of the valence and conduction bands at the Γ point. Analysis of the HSE06 calculated total and partial electronic density of states (PEDOS), Fig. 1(a), shows that the upper valence band is dominated by O 2p states, with the lower conduc-

<table>
<thead>
<tr>
<th>PBE</th>
<th>HSE06</th>
<th>Experiment</th>
</tr>
</thead>
<tbody>
<tr>
<td>a</td>
<td>5.0602</td>
<td>4.9602</td>
</tr>
<tr>
<td>c</td>
<td>3.4546</td>
<td>3.3759</td>
</tr>
<tr>
<td>E_{dir}</td>
<td>−0.70</td>
<td>0.35</td>
</tr>
<tr>
<td>E_{ind}</td>
<td>⋮</td>
<td>0.23</td>
</tr>
<tr>
<td>VB width</td>
<td>8.02</td>
<td>8.45</td>
</tr>
<tr>
<td>O site occupancy</td>
<td>0.984 (4)</td>
<td>0.984 (4)</td>
</tr>
</tbody>
</table>

*aObtained from refinement of neutron diffraction data of β-PbO2.
tion band comprised mostly of Pb 6s states. The HSE06 calculated band structure is shown in Fig. 1(b). The VBM is situated at the R point, with the conduction band minimum situated at Γ point, resulting in an indirect band gap of 0.23 eV and a direct band gap of 0.35 eV at the Γ point.

To understand the difference between the fundamental indirect band gap, and the optical band gaps reported in the literature [12,44], we have computed the optical absorption spectra for PbO₂, with the results presented in Fig. 1(c). The onset of optical absorption begins at ~1 eV and increases steadily. The reason for the large difference between the optical band gap and the fundamental band gap can be explained by analyzing the allowed transitions from the valence band to the conduction band. The onset of strong optical absorption does not begin until 0.70 eV below the VBM at Γ, as transitions from above this point are symmetry forbidden. This type of behavior is known for isoelectronic SnO₂ [45,46], and similar behavior has been established for In₂O₃ [47], and Ti₂O₃ [48].

The conduction band of PbO₂ is found to be dispersive around the zone center, with an electron effective mass of 0.18mₑ. The calculated mass is significantly smaller than the value of 0.80mₑ derived from electrical measurements of a highly degenerate sample [10]. However, given the known values of ca. 0.3mₑ for ZnO, In₂O₃, and SnO₂, which all have much larger band gaps, a smaller value is expected for PbO₂, and therefore electrical transport measurements for PbO₂ should be revisited.

Defective PbO₂—The formation of charged point defects in a lattice can be compensated by electronic (or ionic [49]) defects. The defects considered in this study include the positively charged V₀ and Pbᵢ, which are electron compensated, as well as the negatively charged oxygen interstitial (Oi) and lead vacancy (VPb) centers, which are hole compensated. The energy of the electrons and holes are determined by the electron chemical potential (Fermi level).

Figure 2 shows a plot of formation energy for all point defects under O-rich and O-poor regimes. For n-type defects, the V₀ is the most stable defect under both sets of conditions for high values of the Fermi level, and will be further stabilized at high temperatures by the configurational entropy of O₂, which is neglected in our model but will be partially compensated by the negative vibrational entropy of the vacancy [50]. The Pbᵢ, which had previously been suggested as the dominant defect [17], is competitive in energy across the span of the band gap, but for highly n-type samples, V₀ will dominate. V₀ exists in the +2 charge state for all chemical potentials: a doubly ionized donor. Our neutron diffraction measurements confirm oxygen stoichiometry at a level of 1.6%, which validates the calculations directly, and is consistent with typical electron carrier concentrations of β-PbO₂.

Based on previous reports of PbO₂ electrodes being lead deficient [14,15,19], one would expect the formation energy of the positively charged defects to be lower than, or equivalent to those of the negatively charged defects. Upon examining Fig. 2, we find that this clearly is not the case. Even under O-rich conditions, the formation energy of 7 eV is so prohibitively high that the concentration of VPb should be negligible under equilibrium conditions. In SnO₂, a similar conclusion is reached about the concentration of VSn [20]. The presence of these defects could be explained by defect aggregates formed in the presence of hydroxylated anion sites. Turning to the Oᵢ, we find that it relaxes from the ideal interstitial site towards a lattice oxygen, displacing it to form a peroxide (O=O dumbell-like) species. This type of behavior has also been noted for CdO [51], ZnO [52,53], Al₂O₃ [54], TiO₂ [55,56], and SnO₂ [20,21]. Although the formation energy of Oᵢ is lower than that of VPb, its ionization levels are deep in the conduction band, indicating that it will not act as an effective charge compensating defect in PbO₂. Under conditions that allow for oxygen exchange with the environment, V₀ will be the dominant donor species, and the material will be substoichiometric on the oxygen sublattice, with charge compensation occurring through electrons in the conduction band.

FIG. 1 (color online). HSE06 calculated (a) electronic density of states for PbO₂ with a gaussian smearing of 0.2 eV, (b) band structure for PbO₂, with the position of the calculated charge neutrality level indicated by the horizontal dashed (green) line, and (c) theoretical optical absorption onset of PbO₂.

246402-3
FIG. 2 (color online). Formation energies for intrinsic defects under O-rich conditions (left panel) and O-poor conditions (right panel).

Doping limits.—Typical carrier densities of as-prepared samples of PbO$_2$ exceed the threshold of the Mott criterion [57,58], such that the material falls into the regime of a degenerate semiconductor, with the Fermi level situated within the conduction band. Utilizing the critical carrier concentration \( n^{1/3} a_0 > 0.26 \), where \( a_0 \) is the effective Bohr radius \( (\frac{4\pi\varepsilon_0\varepsilon_n}{m^*e^2})^{1/2} \), results in an electron radius of 29 Å and an associated critical carrier concentration of \( 7 \times 10^{17} \text{ cm}^{-3} \). Once this threshold is exceeded, the material will behave as a metallic conductor and occupation of the conduction band will result in a Moss-Burstein blue-shift in the onset of optical absorption [59].

We can estimate the doping limit (an upper threshold for equilibrium carrier concentration) following the model of the charge neutrality level [60,61]. Using the formalism proposed by Bechstedt and coworkers [62], we find that the charge neutrality level of PbO$_2$ lies 2.66 eV above the conduction band minimum, indicating that this material will be \( n \)-type. This is consistent with the calculated defect energies, which indicate that no charge compensation by ionic defects will occur until the Fermi level is more than 2 eV above the conduction band. This yields an approximate maximum carrier concentration of \( 10^{21} \text{ cm}^{-3} \), which is in direct agreement with the reported value [10], as well as the high level of oxygen nonstoichiometry observed in the diffraction measurements. Assuming that oxygen vacancy states are doubly ionized an oxygen deficiency of 1.6% as found by neutron diffraction corresponds to a carrier concentration of \( 1.5 \times 10^{21} \text{ cm}^{-3} \).

Summary.—We have reported that PbO$_2$ is a semiconducting material with a finite fundamental band gap. Under standard preparative conditions, the material is grossly (1.6%) oxygen deficient, resulting in high electron carrier concentrations. Remarkably, the material can sustain a high donor concentration without self-compensation by negatively charged point defects, which is further validated by a calculated charge neutrality level that is deep in the conduction band.

This work was supported by SFI through the PI program (Grant No. 06/IN.1/I92 and No. 06/IN.1/I92/EC07). Calculations were performed on the TCHPC computing clusters, the Stokes cluster as maintained by ICHEC, and the University of Bath’s High Performance Computing Facility. A.W. acknowledges support from the U.K.’s HPC Materials Chemistry Consortium, which is funded by EPSRC Grant No. EP/F067496, and the EPSRC solar fuel Grant No. EP/I01330X.