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Article

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A Cationic Diode Based on Asymmetric Nafion[®] Film Deposits

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ABSTRACT: A thin film of Nafion[®], of approximately 5 μm thickness, asymmetrically deposited onto a 6 μm thick film of poly(ethylene terephthalate) (PET) fabricated with a 5, 10, 20, or 40 μm microhole, is shown to exhibit prominent ionic diode behaviour involving cation charge carrier (“cationic diode”). The phenomenon is characterized via voltammetric, chronoamperometric, and impedance methods. Phenomenologically, current rectification effects are comparable to those observed in nano-cone devices where space-charge layer effects dominate. However, for microhole diodes a resistive, a limiting, and an over-limiting potential domain can be identified and concentration polarization in solution is shown to dominate in the closed state.

Introduction

Ion-conducting nano-channels are present in many microporous materials (membranes) and important technologies based on this class of materials are well established. In contrast, novel devices with individual nano-channels have evolved only recently, inspired by important examples in nature,^{1,2} and employed for analytical detection devices^{3,4} or in DNA sequencing tools.⁵ Devices with a single surface coated nano-channel⁶ or a polymer-brush coated nano-cone⁷ have been observed to give current rectification effects. Engineered materials with many of these nano-channels operating in parallel could be employed in new “blue-energy” harvesting/conversion processes,⁸ or in desalination applications.⁹ An underlying feature in these types of technologies/devices is a current rectifying ion channel or an “ionic diode”.¹⁰ Ionic diodes can be considered also as key components in the emerging field of “iontronics”.¹¹

Considerable efforts have been made towards the design of current rectifying ion channels, for example based on “track etch” membranes¹² or glass capillaries prepared with nanoscopic precision.¹³ Surface functionalization or “active gating” can be employed to control surface charges and to embed the diode functionality into the nano-channel structure.¹⁴ Often conical channels¹⁵ are employed to introduce asymmetry into the device and to define a “bottleneck” point where the diode mechanism is active. Hydrogel embedded into conical channels,¹⁶ also macroscopic hydrogel-hydrogel junctions,¹⁷ have been employed in current rectification and for “ionic transistors”.^{18,19} Most of these devices are sensitive to ionic strength with higher ionic strength reducing the extent of the double layer and thereby diminishing the magnitude of ion flow rectification phenomenon. We have recently demonstrated that

polymer materials with an intrinsically microporous structure are able to give pronounced diode effects when deposited “asymmetrically” onto a microhole. For example, polymers of intrinsic microporosity (or PIMs²⁰) have been shown to exhibit “ionic diode” and “ionic flip-flop” effects.²¹ In addition, thin films (300 nm) of PIMs have been shown to give strong and pH-switchable ionic diode effects.²² Similarly, the organometallic framework material ZIF-8 has been suggested for ionic diode applications.²³ Here, we report that a commercial fluoropolymer (Nafion[®]), an intrinsically porous ion-channel containing material, provides excellent current rectification and ionic diode characteristics when fabricated into the appropriate micro-structure. A mechanism responsible for the “open” and “closed” diode is proposed.

Nafion[®] (CAS number 66796-30-3) is a well-known commercial ionomer membrane material^{24,25} with applications in fuel cells,²⁶ in catalysis,²⁷ and in electrolyzers.²⁸ Nafion[®] self-assembles into a “channel structure”^{29,30} that allows cation transport through sulfonate-lined hydrophilic channels (Figure 1) within a hydrophobic fluorocarbon matrix.^{31,32} The negatively charged channels allow permanent uptake and conduction of guest molecules, in particular hydrophobic cations.^{33,34}

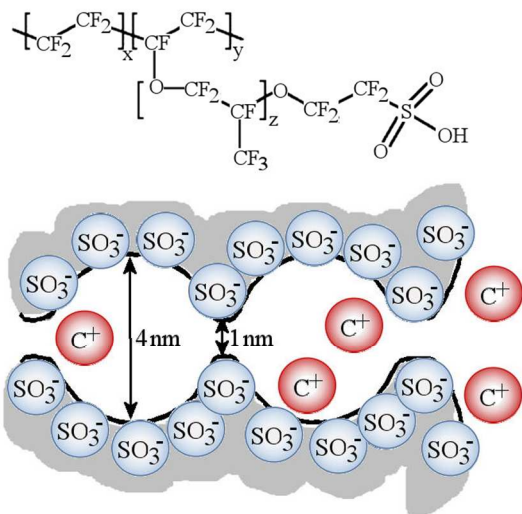


Figure 1. Molecular structure of Nafion[®] and schematic representation of the cation (C⁺) transport channel formed via self-assembly.

In this report, Nafion[®] is applied as a thin (approximately 5 μm thick) film onto a micro-hole in a poly-ethylene-terephthalate (PET) substrate. The voltammetric characteristics of the resulting “asymmetric” device are compared to the corresponding “symmetric” case with Nafion[®] applied to both sides of the PET substrate. It is demonstrated that only the “asymmetric” case results in current rectification or ionic diode phenomena. More generally, it is shown that a *macroscopic asymmetry* in the membrane/aqueous electrolyte interface on opposite sides of the membrane introduces effects that are usually associated with *microscopically asymmetric* nano-cone devices. The physical reasons for the rectification effects are associated with ionomer conductivity in the “open” diode state and with concentration polarization (defined here as a diffusional transport overpotential due to a concentration gradient in the electrolyte) in the external electrolyte close to the PET microhole for the “closed” diode state.

Experimental

Chemical Reagents. Nafion[®]-117 (5 wt. % in a mixture of lower aliphatic alcohols and water), concentrated hydrochloric acid (37%), sodium hydroxide ($\geq 98\%$), concentrated nitric acid (69.0%), concentrated perchloric acid (70%), sodium chloride, potassium chloride, and rhodamine B (97%) were obtained from Sigma-Aldrich or Fisher Scientific and used without further purification. Solutions were prepared under ambient conditions in volumetric flasks with ultra-pure water of resistivity 18.2 M Ω cm from an ELGA Purelab Classic system.

Instrumentation. Electrochemical data (for both voltammetry and impedance) were recorded at $T = 20 \pm 2^\circ\text{C}$ on a potentiostated system (Ivium Compactstat). A classic 4-electrode electrochemical cell similar to that employed in previous membrane conductivity studies³⁵ was used. The membrane

separates two tubular half-cells (15 mm diameter), one with the Pt wire working and saturated calomel (SCE) sense electrode (tied together) and the other with the SCE reference and Pt wire counter electrodes (tied together) (see scheme in Figure 2). The cell when assembled has the shape of a “U” and is therefore referred to as “U-cell”. Fluorescence imaging experiments were performed on a Nikon Eclipse 90i. For fluorescence analysis, rhodamine B was mixed with Nafion[®]-117 (5 wt. %) solution to make a solution (approximately 10 μM of rhodamine B), which was then applied to the PET films.

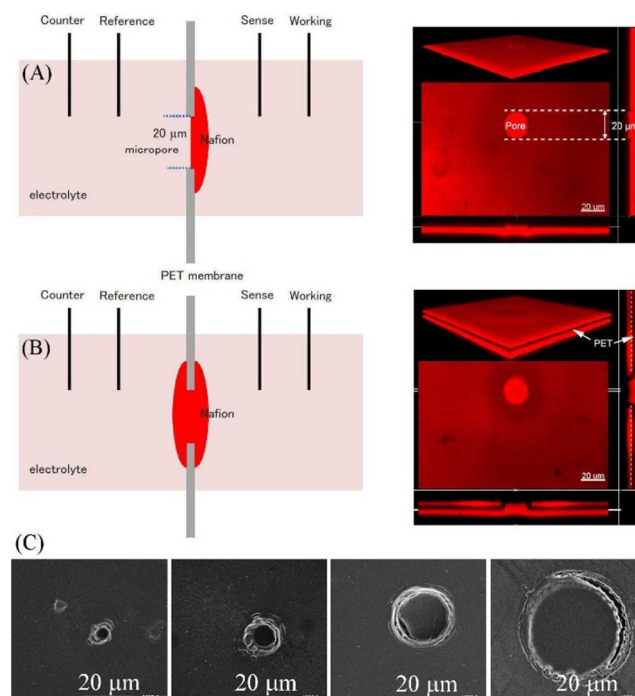


Figure 2. Schematic drawing (not to scale) of the experimental 4-electrode arrangement and fluorescence images for (A) single-sided/asymmetric or (B) double-sided/symmetric Nafion coated (and rhodamine B stained) PET film with 20 μm microhole. Differences in cross-sectional view in top and bottom layer appearance are believed to be linked to light distribution artifacts. (C) SEM images showing PET side for asymmetric Nafion deposits on 40 μm , 20 μm , 10 μm , and 5 μm microholes.

Procedure. In order to form films of Nafion[®] on PET substrates (obtained with 5 μm , 10 μm , 20 μm , 40 μm diameter hole in 6 μm thick PET from Laser-Micro-Machining Ltd., Birmingham, UK), 10 μL Nafion solution was applied to a PET film on a glass substrate (pre-coated with a thin layer of 1% agarose gel to avoid Nafion passing through the microhole) by solution-casting. With a glass rod the Nafion[®] solution was spread evenly over the PET to give a 1 cm^2 film, which after drying produced a thin uniform coating of typically 5 μm thickness. When peeled off the substrate, asymmetric Nafion[®] deposits on the PET microhole are achieved (see electron micrographs in Figure 2C). For symmetric deposits the deposition was repeated on the

opposite side. In order to image the Nafion film fluorescence image stacks were obtained (see Figure 2) for an asymmetric (one-sided) deposit (Figure 2A) and for a symmetric (double-sided) deposit (Figure 2B). In electrochemical measurements with asymmetric deposits the working electrode was always located on the side of the Nafion® film.

Results and Discussion

Nafion® Film Microhole Electrochemistry I: Film Characterization. Fluorescence microscopy with rhodamine B stain (Figure 2) reveals the presence of a uniform $\sim 5 \mu\text{m}$ thick film of Nafion coated over a $20 \mu\text{m}$ diameter microhole in a PET substrate. When the Nafion® is applied in two steps from both sides a “sandwich-like” structure is obtained interconnected through the microhole. These sandwich-films are then placed between two electrochemical half-cells (Figure 2) to allow 4-electrode voltammetric measurements.

Figure 3 shows cyclic voltammetry data for the Nafion® films. Three cases are shown: (i) PET film with an empty $20 \mu\text{m}$ diameter hole, (ii) a similar film but with Nafion® deposited on both sides, and (iii) a similar film but with Nafion® deposited only on one side. For data represented in Figure 3A, aqueous 10 mM HCl solutions are present on both sides of the cell. In the absence of Nafion® (black line) the currents are dominated by the specific resistivity of the aqueous electrolyte solution filling the pore and a typical “Ohmic” slope with $R = 150 \text{ k}\Omega$ is recorded. With the help of equation 1 for left access, transit, and right access resistivity for a microhole,^{36,37} this Ohmic slope can be reconciled with the known specific resistivity of aqueous 10 mM HCl, $\kappa = 0.412 \Omega^{-1}\text{m}^{-1}$.³⁸

$$R = \frac{1}{\kappa} \left(\frac{1}{4r} + \frac{L}{\pi r^2} + \frac{1}{4r} \right) \quad (1)$$

In this equation the microhole resistance R is given by the microhole radius, r , the microhole length, L , and the specific conductivity of the electrolyte, κ .

With Nafion® applied on both sides (blue line) the current clearly increases due to the locally higher concentration of ionic charge carriers (i.e. protons in Nafion®) within the microhole region. The shape of the voltammetric response is again symmetric and dominated by an “Ohmic” slope. With equation 1 the specific resistivity for Nafion® can be estimated as $\kappa = 3.6 \Omega^{-1}\text{m}^{-1}$, which seems realistic.³⁹

Perhaps interestingly, for the asymmetric case (Figure 3A red line) “ionic diode” or current rectification behavior is observed with a higher current at positive bias (consistent with resistive behavior) and a lower current at negative bias, as compared to the empty hole response (consistent with a limiting behavior). Figure 3B shows data for the asymmetric deposit as a function of the electrolyte concentration. Both, currents in the “closed” state and currents in the “open” state increase with ionic strength. The effects on the near-steady state (or time-independent) current signals due to ionic strength and of microhole size are reported in more detail below.

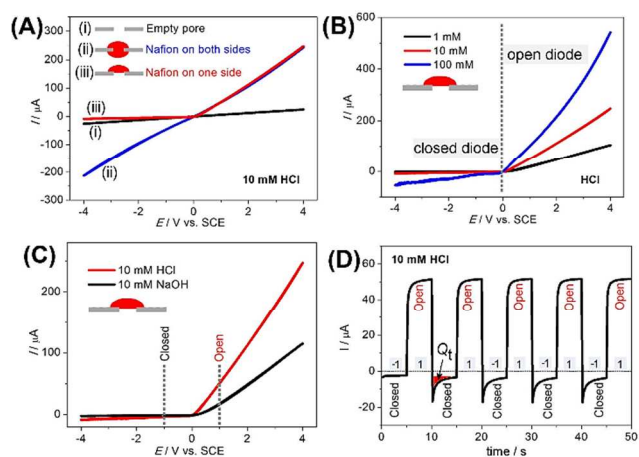


Figure 3. (A) Cyclic voltammograms (scan rate 50 mVs^{-1}) for PET film with a $20 \mu\text{m}$ diameter microhole in a “U-cell” containing aqueous 10 mM HCl on both sides of the PET film showing results for (i) an empty pore, (ii) symmetrically deposited Nafion®, and (iii) asymmetrically deposited Nafion® film. (B) Cyclic voltammograms for an asymmetric Nafion® film immersed in 1 mM HCl, 10 mM HCl, and 100 mM HCl on both sides of the PET film with rectification ratios 13, 15, 7, respectively. (C) Cyclic voltammograms for an asymmetric Nafion® film immersed in 10 mM HCl and in 10 mM NaOH with rectification ratio 15 and 13, respectively. (D) Chronoamperometry data for an asymmetrically deposited Nafion® film immersed in 10 mM HCl when switching the potential between -1 V to $+1 \text{ V}$.

In Figure 3C, cyclic voltammetry data are shown for both 10 mM HCl and 10 mM NaOH environments. Consistently, ionic diode effects are observed with the “open” state always in the positive potential range. This observation is related to the Nafion® structure (in contrast to recent reports of “switching” diode polarity with pH²²) and ascribed here to the fact that Nafion® remains a cation conductor in aqueous HCl (proton transport through Nafion®) and in aqueous NaOH (sodium cation transport through Nafion®). For both, protons or Na^+ cations, similar effects arise at the Nafion® | electrolyte interface with concentration polarization and cation space charge layer effects at the smaller microhole interface linked to the current rectification effect. The higher currents seen for the experiment in 10 mM HCl (compared to currents for 10 mM NaOH) can be explained with the higher proton mobility in Nafion®.⁴⁰

With a scan rate of 50 mVs^{-1} , cyclic voltammetry responses appear to be dominated by near-steady state behaviour, but there are underlying transient processes that can be revealed, for example, by chronoamperometry. Figure 3D shows typical open/closed transients when switching the ionic diode between $+1.0 \text{ V}$ and -1.0 V . Transient responses with a typical time constant of approximately 0.1 s (and underlying charge of typically $Q_t = 20 \mu\text{C}$, see Figure 3D) are observed for both diode opening and diode closing processes. It seems likely that the transient component of the current response is mainly associated with transport of cations, which could imply formation of either a “polymer space charge region” or a “solution concentration polarization region” resulting in the closed

diode state. Nafion[®] has a density of approximately 1.67 g cm^{-3} ⁴⁰ and a molecular mass of approximately 1100 g per mol of sulfonate functional group. For a “plug” of $20 \mu\text{m}$ diameter and $5 \mu\text{m}$ length this amounts to a weight of 2.6 ng and therefore a charge of 2.4 pmol ($= 0.23 \mu\text{C}$). The observed transient charge during switching is two orders of magnitude higher and therefore more likely to be associated with concentration polarization in solution. For a microelectrode concentration polarization should result in an estimated time constant of roughly $\tau = r^2/D$ or here approximately 0.1 s , consistent with the experiment.

An alternative approach to the characterization of transient phenomena is based on electrochemical impedance spectroscopy. Figure 4 shows impedance data for (A) the PET film with empty microhole, (B) the symmetric Nafion[®] deposit, and (C) the asymmetric Nafion[®] deposit. For the empty microhole, a voltage-independent spectrum is obtained with a classic semi-circle indicative of RC-parallel behaviour. In fact the equivalent circuit shown as inset in Figure 4A can be employed to fit the data to give $R1 = 1.3 \text{ k}\Omega$ (consistent with bulk solution resistance), $R2 = 148 \text{ k}\Omega$ (consistent with only microhole and access region resistance, see equation 1), and $C = 0.8 \text{ nF}$ (consistent with PET film capacitance).

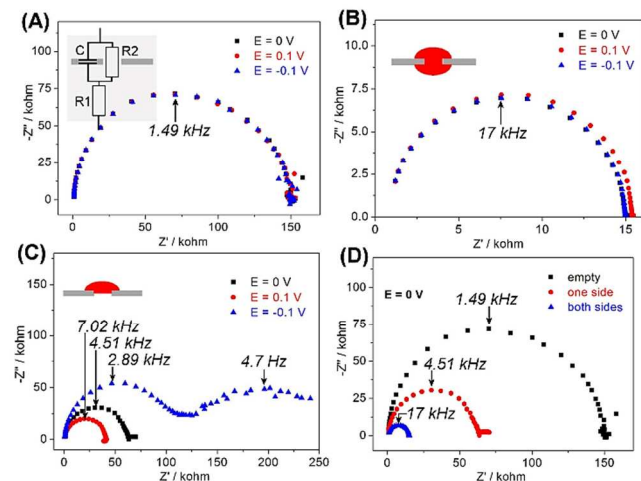


Figure 4. Electrochemical impedance data (frequency range 100 kHz to 1 Hz ; amplitude) for a $20 \mu\text{m}$ diameter pore in PET immersed into 10 mM HCl on both sides compared to data obtained with symmetric and with asymmetric Nafion[®] deposits. Data are shown for (A) the PET microhole without Nafion[®], (B) the microhole with Nafion[®] applied to both sides, (C) the microhole with Nafion[®] applied to only one side, and (D) a comparison of empty, one sided, and two-sided Nafion[®] for an applied voltage of 0 V .

Figure 4B shows impedance spectroscopy data for the PET microhole coated symmetrically with Nafion[®]. Again a potential independent impedance is observed with give $R1 = 0.9 \text{ k}\Omega$ consistent with bulk resistance for both aqueous solution and the polymer region outside the access region), $R2 = 14 \text{ k}\Omega$ (consistent with microhole and access Nafion[®] resistance with-

in the polymer), and $C = 0.7 \text{ nF}$ (consistent with PET film capacitance). Figure 4C shows data for the asymmetrically deposited Nafion[®] film and for this case clear changes in impedance as a function of applied potential are noted. The high frequency semicircle remains (for $-0.1/0.0/0.1 \text{ V}$ applied potential, respectively) with $R1 = 0.18/0.83/1.3 \text{ k}\Omega$ and $C = 0.6 \text{ nF}$ at all three potentials. Parameter $R2 = 99/63/40 \text{ k}\Omega$ mostly reflects the internal resistance in the Nafion[®] film, which appears to go down when the diode “opens”. Intriguingly, a second non-ideal semicircle is observed (Figure 4C) with a time constant of approximately 0.1 s order of magnitude, consistent with transient phenomena during opening/closing processes of the ionic diode. This time constant is proposed to be again associated with concentration polarization.

Nafion[®] Film Microhole Electrochemistry II: Anion versus Cation Effects. The effect of anions and cations are investigated, as well as ionic strength imbalances in left and right half-cells, in order to better understand “cationic diode” processes of the Nafion[®] asymmetrically deposited onto microhole-containing PET film. Figure 5A shows cyclic voltammetry data for aqueous 10 mM acids: HCl , HNO_3 , and HClO_4 . For the asymmetrically deposited Nafion[®] film, reproducible diodes (current rectifiers) are observed in all three cases. The effect of chloride, nitrate, or perchlorate anions remain insignificant.

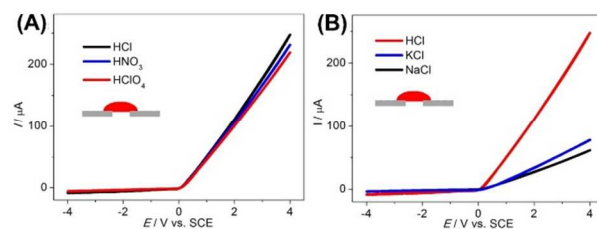


Figure 5. (A) Cyclic voltammograms (scan rate 50 mVs^{-1}) for an asymmetric Nafion[®] membrane immersed on both sides with aqueous 10 mM HCl , HNO_3 , or HClO_4 with rectification ratios $15, 19, 19$, respectively. (B) As above, but with 10 mM HCl , KCl , and NaCl with rectification ratios $15, 9, 13$, respectively.

In contrast, when investigating the effect of the electrolyte cation, much more obvious changes in diode behavior are observed (Figure 5B). Currents for the open diode are considerably higher in the presence of protons compared to those for Na^+ or K^+ . This is consistent with a higher mobility for protons relative to Na^+ and K^+ . Values for ion diffusivity within Nafion[®]⁴¹ in the presence of 1 mol dm^{-3} aqueous chloride solution have been reported as $D(\text{H}^+) = 5.3 \times 10^{-10} \text{ m}^2 \text{ s}^{-1}$, $D(\text{Na}^+) = 1.58 \times 10^{-10} \text{ m}^2 \text{ s}^{-1}$, $D(\text{K}^+) = 0.86 \times 10^{-10} \text{ m}^2 \text{ s}^{-1}$, and for lower salt concentrations $D(\text{Na}^+) \approx D(\text{K}^+)$. This trend is in good agreement with the current – potential data shown in Figure 5B. The mobility of cations in the Nafion[®] film deposit appears to be crucial in determining the magnitude of the current in the open diode state.

When changing asymmetrically the supporting electrolyte concentration, there are two cases to consider. Figure 6A

shows data for cyclic voltammograms employing an aqueous 1 mM HCl solution on the side of the Nafion[®] membrane whilst changing the aqueous electrolyte from 1 mM, 10 mM, 100 mM, to 500 mM on the opposite side of the PET microhole. It can be observed that the closed state of the diode is strongly affected with almost complete “opening” at 500 mM HCl. The current response in the “open” state of the diode remains largely unaffected. It can be concluded that the “closed” state of the diode is due mainly to a region close to the small area Nafion[®] | electrolyte interface in the vicinity of the PET microhole. This region can be identified as the concentration polarization region in the electrolyte solution close to the microhole. Yossifon and coworkers⁴²⁻⁴⁵ and Chang⁴⁶ have shown that for nano-slot channels external concentration polarization causes a “limiting” region where current is limited by an electroneutral diffusion-migration layer similar to that observed at micro-electrodes under conditions of no added supporting electrolyte.⁴⁷ This can be compared to the concentration polarization case here. For HCl as a 1:1 electrolyte it is possible to contrast to the literature concentration polarization case of 1:1 ferricinium salt reduction at a metal microelectrode without added electrolyte. At the metal electrode the diffusion-migration transport is linked to electron transfer, whereas at the diode the diffusion-migration transport is linked to cation flux through the Nafion[®] layer. The expected mass transport limited current⁴⁸ is $I_{lim} = Z \times 4 F D r c$ (with $Z = 2$ for a 1:1 electrolyte, F , the Faraday constant, D the diffusion coefficient, r the microdisc radius, and c the concentration). Assuming $D_{proton} = 9 \times 10^{-9} \text{ m}^2 \text{ s}^{-1}$, the estimated limiting current is $I_{lim} = 35 \text{ } \mu\text{A}$, which is not too far from the observed value $I_{lim} = 50 \text{ } \mu\text{A}$ (see Figure 6A, “limiting region”). Therefore, following on Yossifon’s work, the “limiting region” as well as the “resistive region” and the “over-limiting region” (caused by convection⁵⁰) can be identified. For Nafion[®] asymmetrically deposited onto a microhole the closed state of the ionic diode is dominated by concentration polarization in solution.

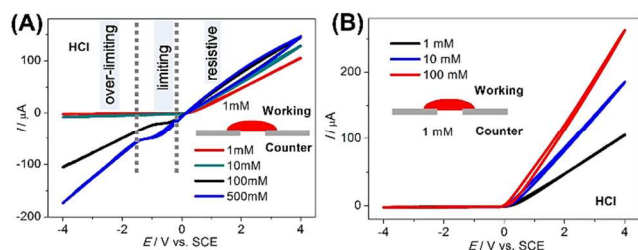


Figure 6. (A) Cyclic voltammograms (scan rate 50 mVs^{-1}) for an asymmetric Nafion[®] membrane immersed on the working electrode sides in 1 mM HCl and on the counter electrode side with 1, 10, 100, 1000 mM HCl. (B) As above, but with 1 mM HCl at the counter electrode side and 1, 10, 100 mM (with rectification ratios 31, 18, 13, respectively) at the working electrode side.

In the case of the asymmetrically deposited Nafion[®], equal concentrations of HCl on both sides of the membrane, and a closed state, the high resistivity in the “limiting” or “overlimiting” regime is therefore determined by the concentration polarization (*i.e.*, cation depletion occurs with a gradient of C^+ and A^- concentrations in the diffusion layer whilst maintaining

electroneutrality) on the microhole side of the Nafion[®] membrane. The reduced conductivity close to the PET microhole caused by concentration polarization (*i.e.*, small number of ions) limits/determines the overall current.

Data shown in Figure 6B demonstrate the complementary case of varying the aqueous electrolyte concentration in the half-cell facing the Nafion[®] film from 1 mM, 10 mM, to 100 mM. In this case the “closed” state of the diode (negative potential range) is unaffected and only the “open” state of the diode (positive potential range) is seen to significantly increase in current with increasing HCl concentration. The increase in current is not linear with aqueous solution concentration and may be linked to some additional partitioning of HCl into the Nafion[®] film (affecting the proton mobility in the Nafion[®]). In addition to the nature of cation and anion and the electrolyte concentration effects, it is of interest to explore microhole size/geometry effects on the diode characteristics.

Nafion[®] Film Microhole Electrochemistry III: Microhole Size Effects. Experiments were performed with a $\sim 5 \text{ } \mu\text{m}$ thick Nafion[®] film deposited onto laser-drilled PET films of a systematically increasing diameter (see Figure 2C). Cyclic voltammetry data in Figure 7 demonstrate the diode effect for 10 mM HCl (in both half-cells) for $5 \text{ } \mu\text{m}$, $10 \text{ } \mu\text{m}$, $20 \text{ } \mu\text{m}$, and $40 \text{ } \mu\text{m}$ diameter microholes. As expected, the increase in diameter causes an increase in the currents for both open and closed diode. However, the currents do not scale with area. The relative change in current magnitude for open and closed states shows a pattern indicative of a slightly better current rectification at smaller microholes. The diameter of the microhole is an important parameter in controlling ionic diode characteristics, but other parameters such as Nafion[®] film thickness and PET film thickness may be equally important and will require further study.

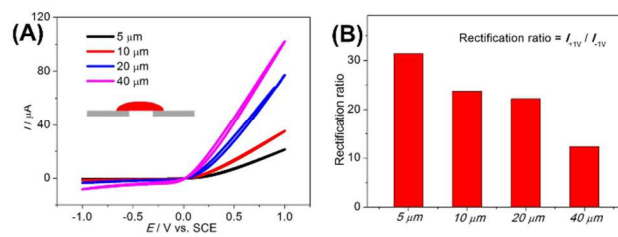


Figure 7. (A) Cyclic voltammograms (scan rate 50 mV s^{-1}) in 10 mM HCl and for 5, 10, 20, and $40 \text{ } \mu\text{m}$ diameter microhole in a $6 \text{ } \mu\text{m}$ thick PET film coated asymmetrically with Nafion[®]. (B) Bar graph for the current rectification ratio at $\pm 1 \text{ V}$ versus microhole diameter.

Comparing the present data for Nafion[®] on a microhole to previous studies of nano-cone ion current rectification processes is possible, if we assume Nafion[®] to be composed of many nano-channels. Asymmetry for the nano-cone is introduced at microscopic level at the interface to the electrolyte solution. Asymmetry for the Nafion[®] films is introduced at macroscopic level due to the difference in area exposed to the opposing electrolyte solutions and the difference in diffusion-

1 migration access to the microhole. Nano-cone effects are
 2 based on double layer phenomena (formation of a space
 3 charge layer or accumulation-depletion⁵¹), which also lead to
 4 changes in the apparent activation energy for ion transport⁵² in
 5 these devices. The accumulation-depletion model has been
 6 developed by White and coworkers^{53,54} and has been employed
 7 to account for current rectification in single cone nano-
 8 channels. A variation of the ion concentration in the vicinity of
 9 the cone-tip to solution phase interface has been identified as
 10 the prime reason for local changes in ion conductivity that
 11 give rise to the current rectification effect. Here, for Nafion[®],
 12 the role of the asymmetric cone shape for a single channel is
 13 replaced by the asymmetry in the Nafion[®] microhole deposit
 14 exposed to the electrolyte solution phase. The effect of the
 15 space charge layer is complemented by the concentration polar-
 16 ization phenomenon in solution. In fact, for the highly cation-
 17 conductive Nafion[®] the concentration polarization phenom-
 18 enon dominates. In a qualitative manner, the field applied
 19 externally to the Nafion[®] film can be suggested to:

- 20 (A) drop across the whole Nafion[®] film for the “*open di-*
 21 *ode*” with film resistivity (mainly in the access re-
 22 sistance region close to and within the microhole)
 23 dominating the current flow and
 24 (B) drop primarily across the small Nafion[®] | electrolyte
 25 interface within the microhole for the “*closed diode*”
 26 associated with a space charge layer and with concen-
 27 tration polarization in the solution phase, which
 28 leads to loss of electrolyte and conductivity (deple-
 29 tion) locally in the interfacial region;
 30 (C) for the case of Nafion[®] the high ionic conductivity in
 31 the polymer and the high concentration of mobile
 32 cations in the polymer cause concentration polariza-
 33 tion in the electrolyte close to the microhole to domi-
 34 nate the resistivity in the “*closed*” state.

35 Figure 8 shows a schematic drawing to summarize the condi-
 36 tions at the Nafion[®] | electrolyte interface for the closed diode.
 37 The concentration of cations in both adjacent electrolyte solu-
 38 tion and Nafion[®] film approach zero. In the “*over-limiting*”
 39 state additional convective phenomena arise. In the “*resistive*”
 40 or “*open*” state, the Nafion[®] film resistance limits current
 41 flow. In future, further quantitative modelling based on Pois-
 42 son-Nernst-Planck methods (whilst taking into account the ge-
 43 ometry of the Nafion[®] on PET) will be necessary to confirm
 44 these ideas and to allow quantitative prediction of device per-
 45 formance for example in desalination or in energy conversion.

46 Conclusion

47 It has been shown that asymmetry in the deposition of an ion-
 48 nano-channel material, such as commercial Nafion[®] 117, can
 49 be used to induce current rectification phenomena or ionic
 50 diode effects. In this particular case a “*cationic diode*” has
 51 been produced with effective rectification for many types of
 52 cations. The polymer | electrolyte interface in the microhole
 53 provides the key to current rectification effects with currents
 54 for “*open*” diodes being dominated by ionomer conductivity
 55 and currents for “*closed*” diodes being dominated by external
 56
 57
 58
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 60

concentration polarization in the electrolyte. Smaller micro-
 holes have been shown to produce improved current rectifica-
 tion and are likely to allow faster diode switching times.⁵⁵

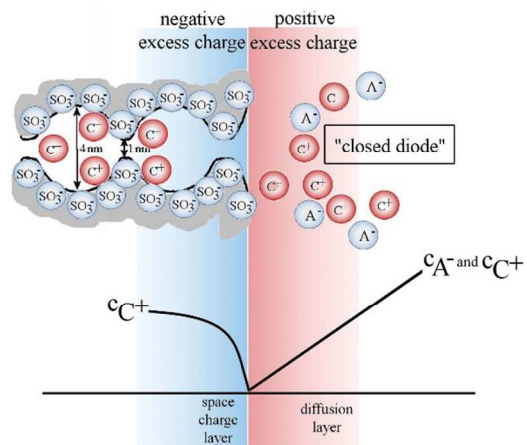


Figure 8. Schematic depiction (not to scale) of the “*closed diode*” Nafion[®] | aqueous electrolyte interface within the microhole and of the effect of the interfacial polarization (presence of excess charge) with concentration polarization in the diffusion layer and cation concentration approaching zero at the interface.

In future, many other semipermeable materials should become available for the development of improved microhole current rectification devices. Both “*cationic diodes*” and “*anionic diodes*” as well as ion-selective diodes are possible. Particularly interesting is the prospect for chemical modification of the ionomer | electrolyte interface to provide switchable or pH sensitive devices.⁵⁶ Applications could be possible in bio-mimetic ion gates,⁵⁷ biosensors,⁵⁸ photo-responsive ion gates,⁵⁹ stimuli-response gates,⁶⁰ or in “*ionic sensors*” which are devoid of immediate metal components and are based entirely on ion conductors in the sensing mechanism.

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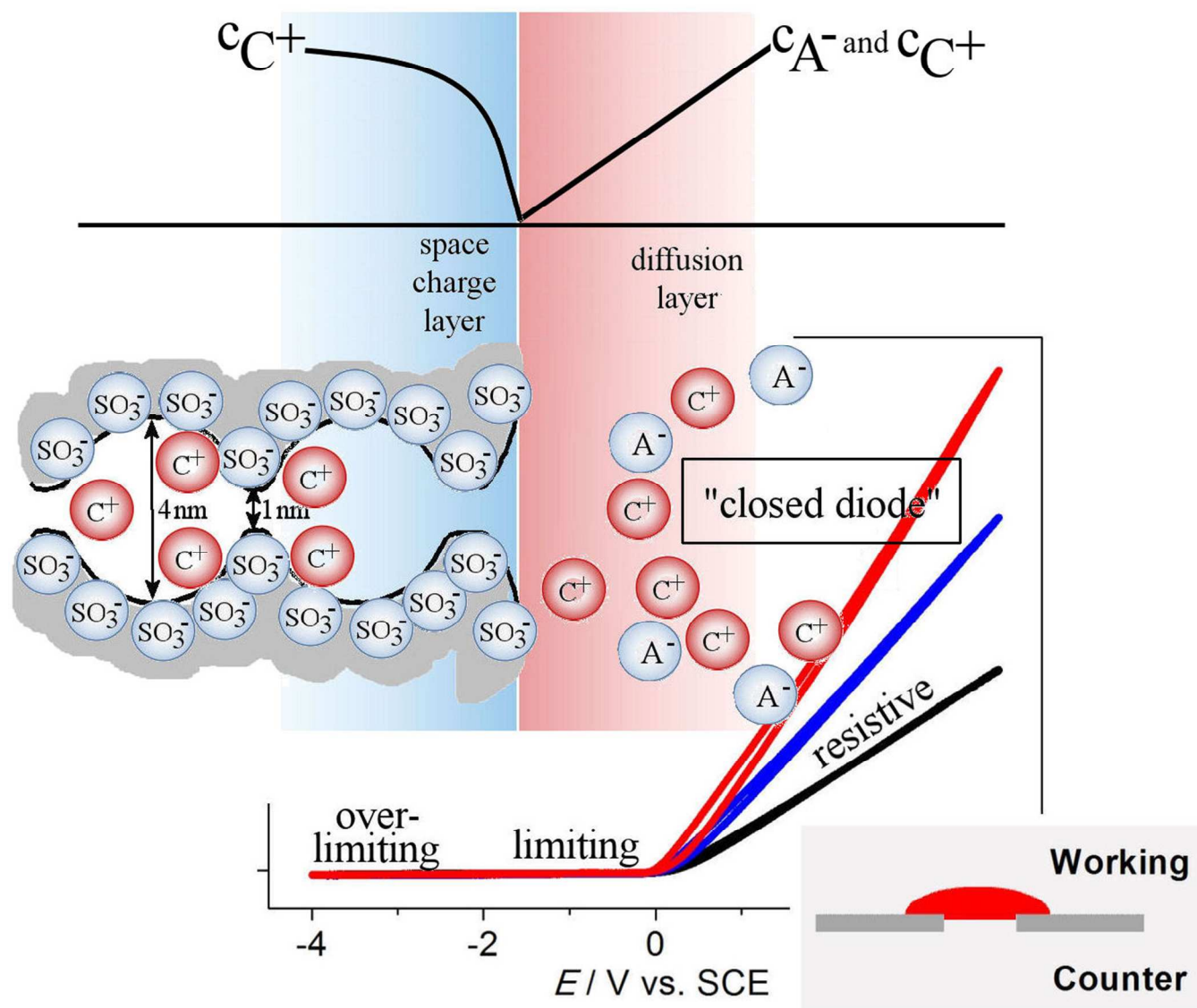
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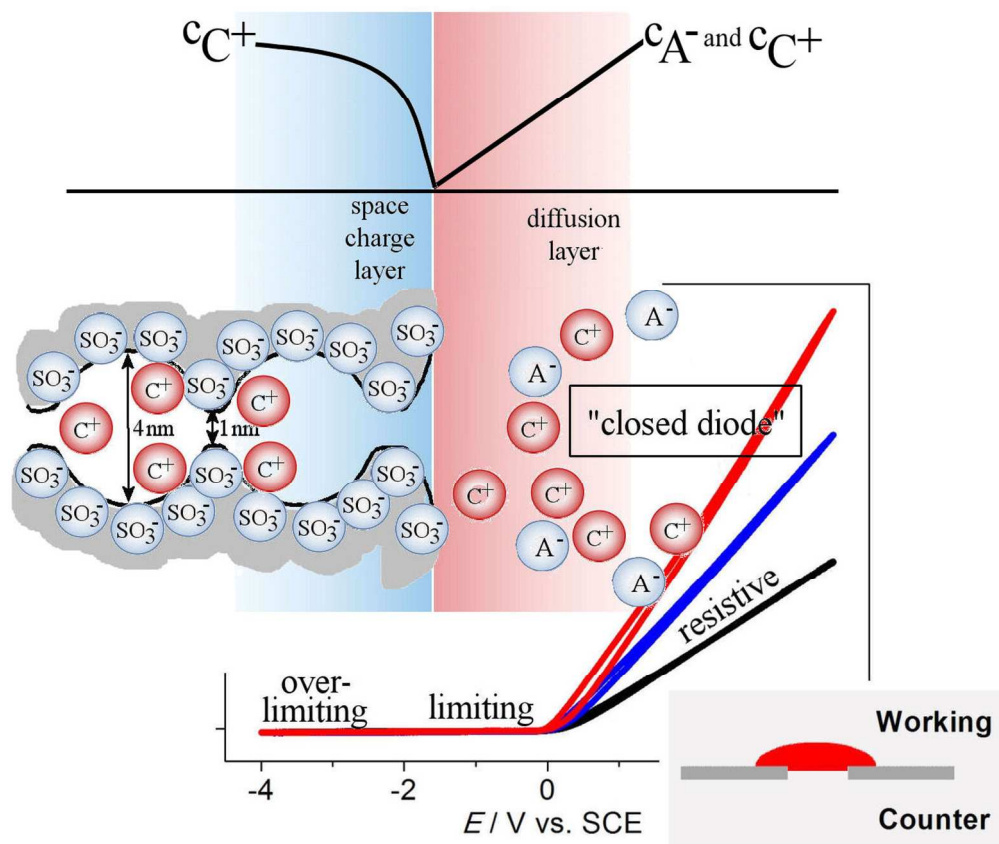
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GRAPHICAL ABSTRACT

**SYNOPSIS for Graphical Abstract**

Nafion[®]-117 cation conductive films are applied about 5 μm thick to a microhole in poly-ethylene-terephthalate (PET) to give pronounced current rectification phenomena that are related to those observed in nanofluidic "ionic diodes", but associated with processes in both polymer and electrolyte phase.



Graphical Abstract

483x417mm (72 x 72 DPI)

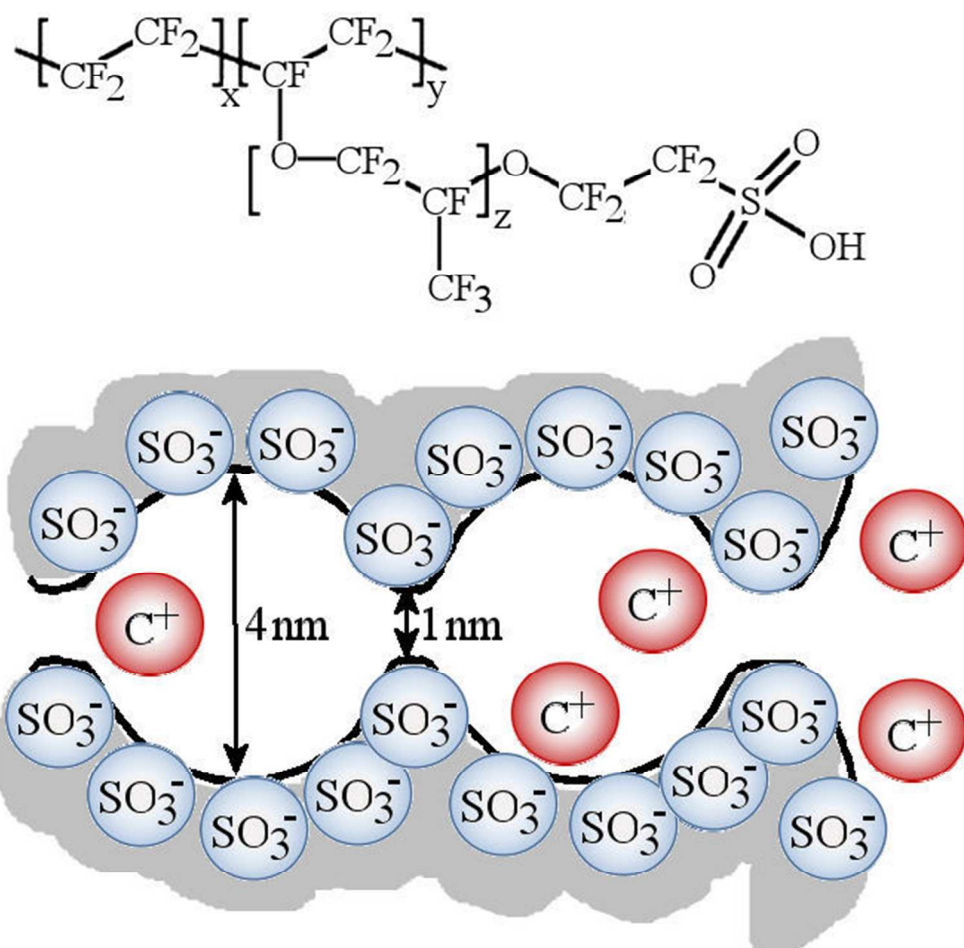


Fig. 1

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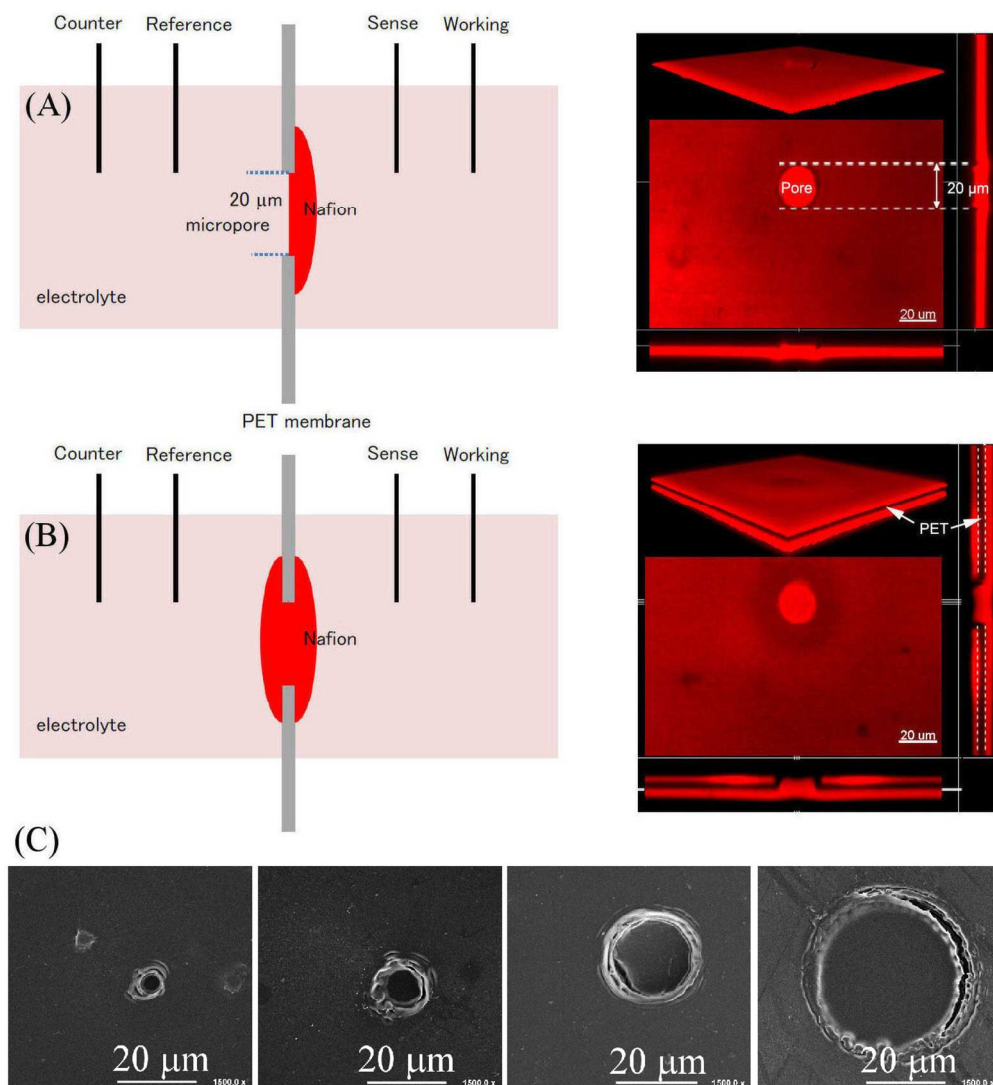


Fig. 2

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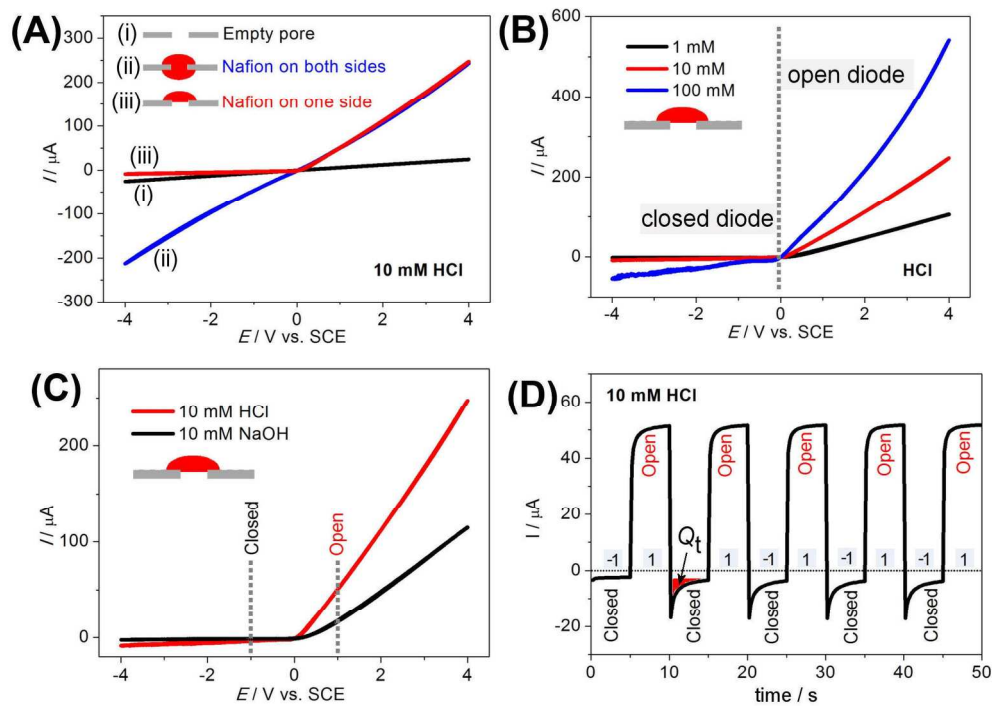


Figure 3

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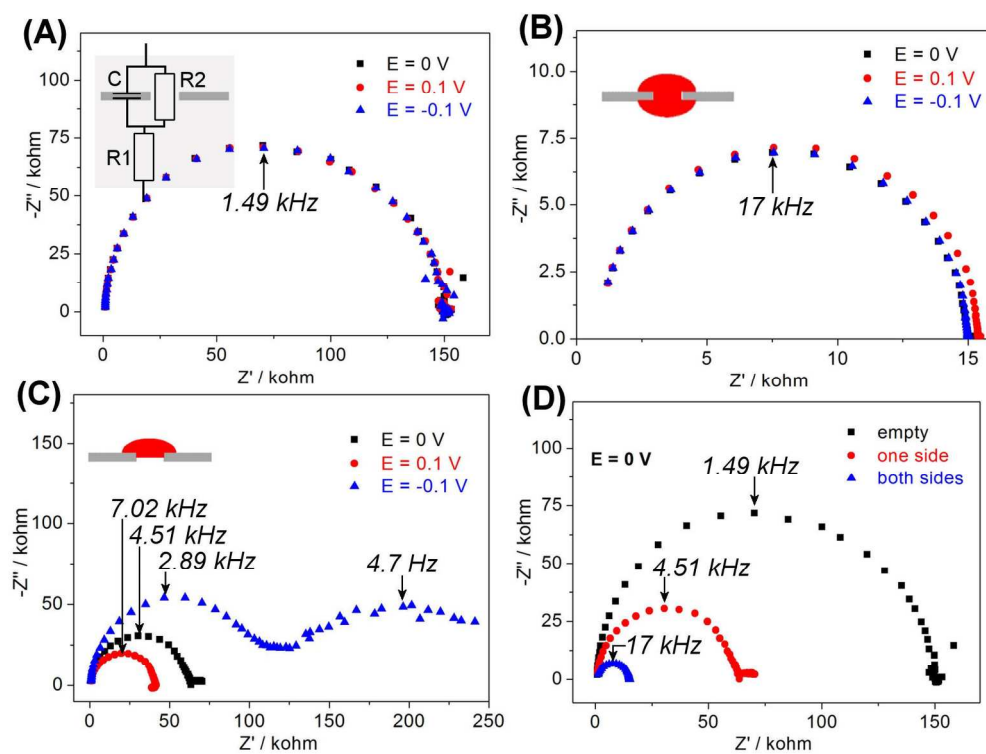


Figure 4

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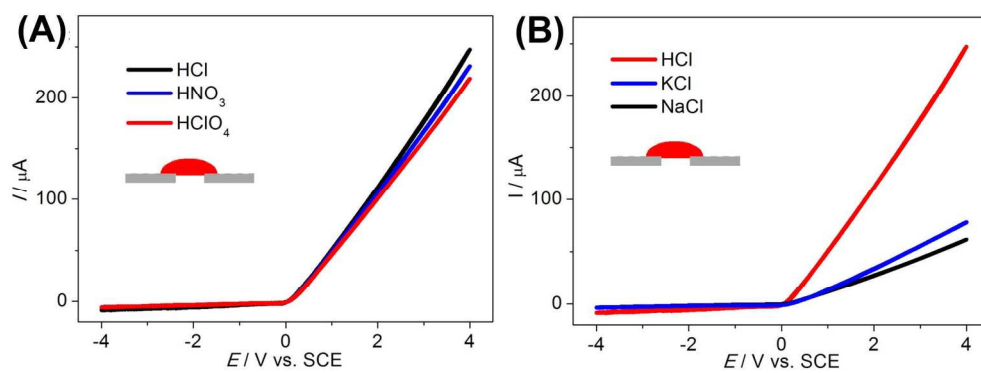


Fig. 5

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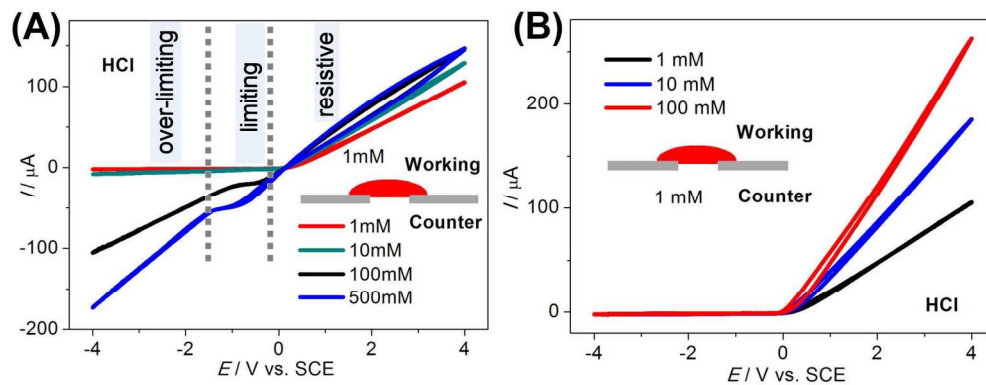
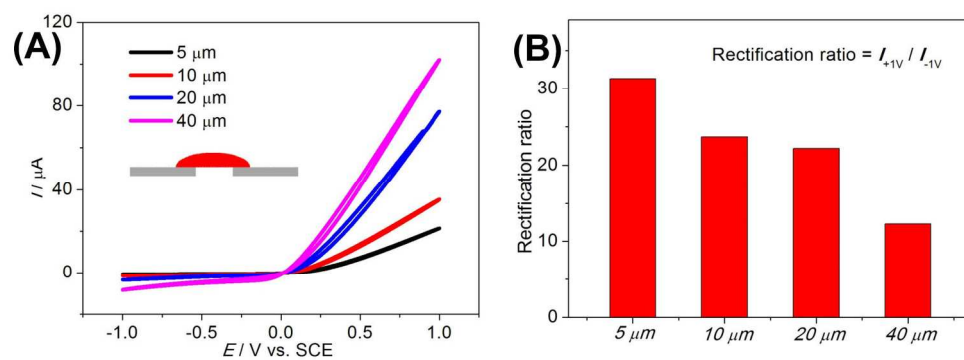
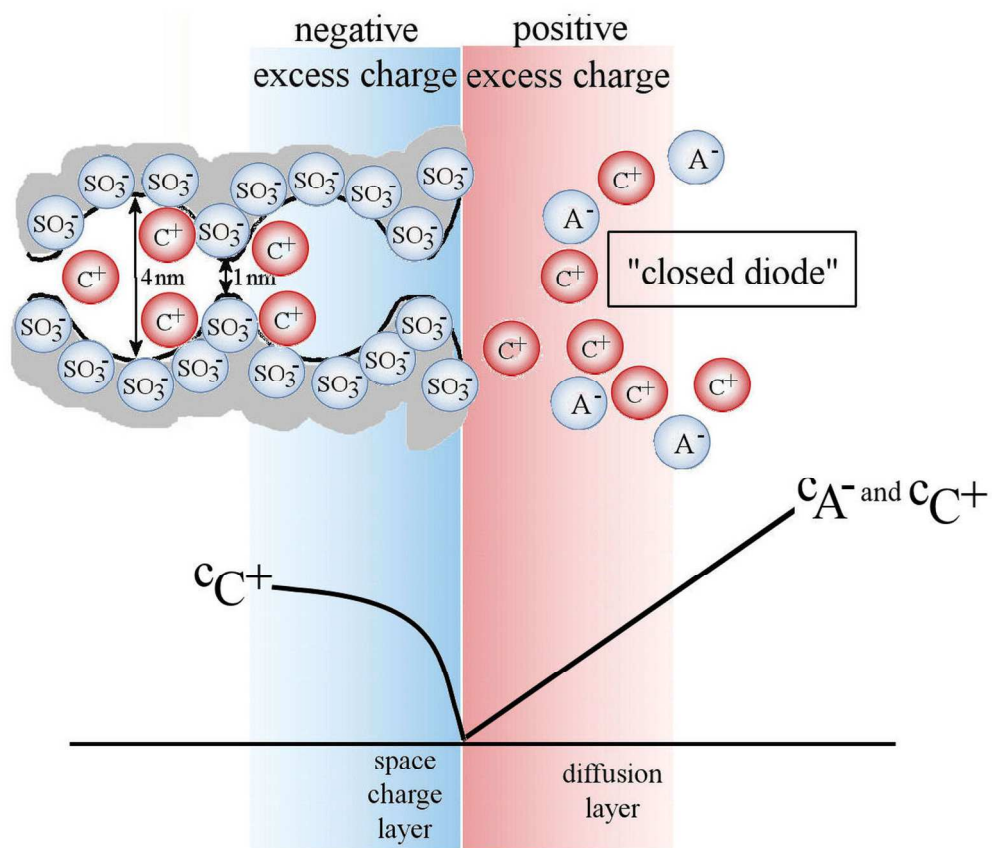


Figure 6

734x282mm (72 x 72 DPI)



745x275mm (72 x 72 DPI)



452x394mm (72 x 72 DPI)