



University of Groningen

## De structuur van I2Cl6(ICl3)

Boswijk, Klaas Harm

IMPORTANT NOTE: You are advised to consult the publisher's version (publisher's PDF) if you wish to cite from it. Please check the document version below.

Document Version Publisher's PDF, also known as Version of record

Publication date: 1954

Link to publication in University of Groningen/UMCG research database

Citation for published version (APA): Boswijk, K. H. (1954). De structuur van I2Cl6(ICl3) Groningen: Excelsior

Copyright Other than for strictly personal use, it is not permitted to download or to forward/distribute the text or part of it without the consent of the author(s) and/or copyright holder(s), unless the work is under an open content license (like Creative Commons).

If you believe that this document breaches copyright please contact us providing details, and we will remove access to the work immediately and investigate your claim.

Downloaded from the University of Groningen/UMCG research database (Pure): http://www.rug.nl/research/portal. For technical reasons the number of authors shown on this cover page is limited to 10 maximum.

## SUMMARY

In the introduction some chemical and physico-chemical properties of the interhalogen compounds are reviewed. Making use of existing thermodynamical data their stability is discussed in some more detail.

The interhalogens, especially the higher ones, stimulated theoretical speculations about their structures. For the greater part the recent structure investigations on these compounds, however, are not in agreement with the theoretically proposed structures.

A contribution to the structure chemistry of the interhalogen compounds has been given by the determination of the crystal structure of ICl, by means of X-ray methods.

Ordinary and integrated Weissenberg photographs from the first nine levels about the needle axis of the crystals, as well as Laue-, oscillation- and flat film photographs were made. The interpretation of the photographs caused some difficulties, which could be overcome by assuming the crystals to consist of four complexly twinned triclinic individuals. The lattice constants are:

a	=	5.71	±	0.03	AO	α =	:	130 <sup>0</sup> 50'	t	10'
b	=	10.88	±	0.02	AO	β =	:	80 <sup>0</sup> 50'	±	20'
с	=	5.48	±	0.03	AO	γ =	:	10 <sup>0</sup> 30'	±	10'

Two ICl, per unit cell; space groep P1.

Because of the complexity of the reciprocal lattice only one - the [010] - projection was accessible for two dimensional Patterson and Fourier methods. The [010] Patterson synthesis was first tentativity interpreted in terms of flat ICl<sub>3</sub> molecules with valence angles of  $120^{\circ}$ . Fourier syntheses based on this structure could not be refined. A correct approximation to the structure was obtained by the application of the vector convergence method on the (101) Patterson section. It could be confirmed by a Fourier syntheses of the (101) section which, however, still showed considerable false detail. The atomic coordinates were refined by a sequence of ordinary and difference Fourier syntheses. The false detail proved to be partly due to incorrect scaling factors relating reflections of different layer lines, and partly to the finite series effect.

The latter could be reduced by the introduction of an artificial temperature factor. The final coordinates, corrected for the remaining effect of the back shift

The final value of The crystal st arranged in layers the crystal may be The molecules a point group symmetried by two chlorine two others at 2.68 The molecular st between valence pat The possibility o ICl<sup>+</sup><sub>2</sub>- and ICl<sup>-</sup><sub>4</sub> ions

70

remaining effect of termination of the series by the application of the back shift method, are:

	х	У	$\mathbf{Z}$
Ι	0.0	67.0	0.0
C1,	99.1	48.6	94.7
C1,	97.9	167.2	92 <b>.</b> 7
$Cl_3$	97.8	286.5	92.0

The final value of the disagreement index R is 0.12.

The crystal structure consists of molecules  $I_2Cl_6$  which are arranged in layers (101). The two types of twinning occurring in the crystal may be explained by this structure.

The molecules are planar with negligible deviations from the point group symmetry mmm (fig. 15). Each iodine atom is surrounded by two chlorine atoms at a distance of 2.38 and 2.39  $A^{O}$  and by two others at 2.68 and 2.72  $A^{O}$ .

The molecular structure is described in terms of a mesomerism between valence patterns involving  $ICl_2^+$  and  $ICl_4^-$  ions (pag. 67). The possibility of an average structure consisting of actual  $ICl_2^+$  and  $ICl_4^-$  ions could be excluded.

d physico-chemical proare reviewed. Making use tability is discussed in

nigher ones, stimulated uctures. For the greater on these compounds, howretically proposed struc-

stry of the interhalogen mination of the crystal ods.

notographs from the first he crystals, as well as tographs were made. The some difficulties, which stals to consist of four s. The lattice constants

 $= 130^{\circ}50^{\circ} \pm 10^{\circ}$  $= 80^{\circ}50^{\circ} \pm 20^{\circ}$  $= 10^{\circ}30^{\circ} \pm 10^{\circ}$ 

Ρ**ī**.

iprocal lattice only one for two dimensional Pat-Patterson synthesis was of flat  $ICl_3$  molecules syntheses based on this ect approximation to the tion of the vector consection. It could be con- $I0\overline{1}$ ) section which, howletail. The atomic coorordinary and difference oved to be partly due to ctions of different layer ffect.

troduction of an artifiinates, corrected for the

71