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# Bond Length – Bond Valence Relationships for Carbon – Carbon and Carbon – Oxygen Bonds



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Running Title: Bond Length – Bond Valence Relationships for C-C and C-O Bonds

## Abstract

In the present study, relationships are developed for determining bond orders (also referred to as bond valences or bond numbers) from published bond lengths for carbon-carbon (C-C) and carbon-oxygen (C-O) bonds. The relationships are based on Pauling's empirical formula  $s = \exp((R_o - R)/b)$ , where  $s$  is the bond order,  $R$  is the corresponding bond length,  $R_o$  is the unit valence bond length, and  $b$  is a fitting parameter. We use a recently derived relationship for the  $b$  parameter in terms of the bonding atoms' published atomic orbital exponents. The resulting equations were checked against published x-ray diffraction (XRD) data for 176 carbon systems with 540 published C-C bond lengths, and 50 oxygen systems having 72 published C-O bond lengths. The C-C and C-O bond length-valence relationships are shown to have sufficient applicability and accuracy for use in any bonding environment, regardless of physical state or oxidation number.

## Introduction

In 1929, Linus Pauling published his five rules of chemical bonding which could be used for predicting crystal structures (Pauling 1929). Pauling's second rule is that of local charge neutrality, commonly known as the valence sum rule, which states that the charge of an anion is neutralized by the sum of the adjacent cationic charges, while any cationic charge is neutralized by adjacent anionic charges. In terms of bond valence, the total valence at any one atom is equal to the sum of that atom's individual bond valences. In 1947, Pauling published the following empirical bond length-valence relationship:

$$s = \exp\left(\frac{R_o - R}{b}\right) \quad (1)$$

where  $s$  is the bond valence, which corresponds to the number of lone pairs of electrons contributing to the bond,  $R_o$  is the length of a chemical bond with unit valence,  $R$  is an observed bond length, and  $b$  is an empirical fitting parameter (Pauling 1947). A wide range of determined values for the  $b$  parameter, anywhere from 0.25 to 0.65 Å (Hardcastle 2013), led to many inconsistencies in valence values, an issue that hindered the ability of chemists to compare findings. As a result, it was proposed that a universal value of 0.37 Å for  $b$  be established as the average from the crystallographic data (Brown and Altermatt 1985). This resulted in consistent relationships having only one fitting parameter,  $R_o$ ; however, when applied to shorter and longer bonds, the calculated valence was shown to be less reliable owing to the inaccuracy of  $b$ .

## Theory

In 2013, Hardcastle derived Pauling's bond length-valence, including a new definition for the  $b$  fitting parameter (Hardcastle 2013). Since then, a slight modification has been made (Hardcastle *unpublished data*), resulting in the following equation:

$$b = \frac{2a_o}{(\xi_1 + \xi_2)} \quad (2)$$

where  $b$  is dependent upon the Bohr radius of a hydrogen atom,  $a_o$  (0.529 Å), and the sum of the orbital exponent values for each of the atoms contributing to the bond. This definition results in values for  $b$  that are specific to the type of chemical bond, a much more accurate alternative to the average universal value of 0.37 Å assumed for any type of bond. Using published values for atomic orbital exponents to determine the value of the  $b$  parameter for any bond, and substituting this value into Equation (1), results in a bond length-valence relationship specific to that bond type.

**Uncertainty in  $R_o$** 

With a definition for the  $b$  parameter,  $R_o$  is left as the only fitting parameter in Equation (1). The precise length of a C-O bond having a bond order of exactly one (a true C-O single bond) is a matter of debate, but has been estimated between 1.33 and 1.43 Å by Allen (Allen et al 1987) and at 1.39 Å by Brese and O’Keeffe, which they refer to as the “bond-valence parameter” because they were using  $b = 0.37$  Å as a universal constant (Brese and O’Keeffe 1991). For the C-C bond length of unit valence, however, most investigators agree on the published C-C length found for crystalline diamond at  $R_o = 1.542$  Å (Brown 2002) as representing the C-C bond length of unit valence.

**Results and Discussion**

The atomic orbital exponents for carbon and oxygen are from data published by Clementi and Raimondi (1963), with values of 1.5679 and 2.2266 respectively. Substituting these values into Equation (2), results in  $b$  parameters of 0.337 Å for C-C bonds and 0.279 Å for C-O bonds. Note that both of these values are much lower than the previously assumed universal constant of 0.37 Å. This leaves  $R_o$ , the bond length of unit valence, as the only remaining fitting parameter. In the case of C-C bonds, our initial guess would be the C-C bond length in diamond at  $R_o = 1.542$  Å. But for C-O bonds, this value could be anywhere from 1.33 (Allen et al 1987) to 1.43 Å (Schomaker and Stevenson 1941).

The total atom valence for a carbon-centered environment was predicted to be 4.00, carbon having four electrons available for bonding, while the atom valence for an oxygen-centered environment was predicted to be 2.00. These predictions were based upon the number of bonding electrons available (oxidation state) in each atom, 4 for carbon and 2 for oxygen. Comparing the calculated atomic valences to the predicted valences, the total error for the C-O bonding was minimized by manipulating  $R_o$ . Clementi’s orbital exponents were not changed to minimize error, but were held constant.

X-ray diffraction data, limited to results published in the year 2000 or later, was collected for C-C and C-O bond lengths, totaling 612 bonds (176 carbon environments and 50 oxygen environments). Each environment is represented by an individual table within Table 1. Bond lengths were recorded and converted to bond valence values, which were then totaled for the atom valence. Data analysis and error

minimization led to two specific relationships, one for C-C bonds:

$$s = \exp\left(\frac{1.5420-R}{0.337}\right) \quad (3)$$

and one for C-O bonds:

$$s = \exp\left(\frac{1.3669-R}{0.279}\right) \quad (4)$$

Each equation was shown to produce accurate valence values from published bond length data. The error in the data (XRD data and valence sum rule) was minimized at  $R_o = 1.5420$  Å for C-C bonds, consistent with the C-C bond length of diamond, and  $R_o = 1.3669$  Å for C-O bonds, consistent with the estimated 1.33-1.44 Å range for a C-O bond of unit valence.

**Conclusion**

Bond length – bond valence relationships, based on Pauling’s formula, provide useful tools for the prediction and evaluation of crystal structures when used with the valence sum rule. In the present study, atomic orbital exponents were used to independently determine the value of the  $b$  parameter (previously either a floating fitting parameter, or set as a universal constant at 0.37 Å) for C-C and C-O bonds at 0.337 and 0.279 Å, respectively. This approach resulted in bond length – valence relationships for C-C and C-O bonds by using published crystallographically determined bond lengths for 612 bonds and the valence sum rule. The optimized bond lengths of unit valence are minimized at  $R_o = 1.5420$  Å for C-C bonds, consistent with the C-C bond length of diamond, and  $R_o = 1.3669$  Å for C-O bonds, consistent with the estimated 1.33-1.44 Å range for a C-O bond of unit valence.

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**Bond Length – Bond Valence Relationships for C-C and C-O Bonds**

Table 1. Bond valence calculations from published XRD data.

<b><i>Carbon Environments</i></b>																		
<b>Alonso et al. 2009</b>			C13	1.338	1.831	C9	Valence	C36	1.403	1.510	C2	Valence						
C10	1.474	1.223	C17	1.508	1.106	C4	1.429	1.398	C43	1.517	1.077	C1	1.423	1.423				
O2	1.210	1.755	O1	1.360	1.025	C8	1.390	1.569			4.035	C3	1.204	2.723				
O3	1.304	1.253			3.962	C10	1.525	1.052			C39	Valence			4.146			
		4.232	C16	Valence				4.019	C34	1.398	1.532			C3	Valence			
C11	Valence		C15	1.341	1.814			C40	Valence	C38	1.398	1.532	C2	1.204	2.723			
C10	1.474	1.223	C21	1.504	1.119	C35	1.410	1.479	C40	1.520	1.067	C4	1.431	1.390				
O2	1.210	1.755	O1	1.361	1.021	C39	1.397	1.537			4.132				4.113			
O3	1.304	1.253			3.955	C41	1.517	1.077	<b>Borbulevych et al. 2002</b>					C6	Valence			
		4.232	C12	Valence				4.093			C8	Valence	C5	1.423	1.423			
C11	Valence		C1	1.393	1.555			C36	Valence	C7	1.500	1.133	C7	1.200	2.756			
C10	1.475	1.220	C13	1.425	1.414	C35	1.413	1.466	C9	1.540	1.006				4.179			
O2	1.213	1.737	C15	1.427	1.406	C37	1.395	1.546	O5	1.432	0.792			C7	Valence			
O3	1.349	1.066			4.376	C44	1.522	1.061	O6	1.430	0.797	C6	1.200	2.756				
		4.023	C3	Valence				4.073			3.728	C8	1.423	1.423				
C11	Valence		C2	1.381	1.612			C4	Valence			C4	Valence			4.179		
C10	1.460	1.275	C4	1.442	1.345	C3	1.426	1.410	C3	1.410	1.479	C3	1.410	1.479			C9	Valence
O2	1.225	1.663	C5	1.430	1.394	C5	1.390	1.569	C5	1.462	1.268	C5	1.462	1.268	C8	1.423	1.423	
O3	1.331	1.137			4.350	C9	1.518	1.074	C9	1.420	1.436	C9	1.420	1.436	C10	1.205	2.715	
		4.076	C6	Valence				4.053			4.182				4.138			
C11	Valence		C5	1.474	1.223			C8	Valence	<b>Burlakov et al. 2015</b>					C10	Valence		
C10	1.468	1.245	C7	1.387	1.583	C3	1.416	1.453			C7	Valence	C9	1.205	2.715			
O1	1.230	1.634	C11	1.383	1.602	C7	1.394	1.551	C6	1.430	1.394	C6	1.430	1.394	C11	1.430	1.394	
O2	1.318	1.192			4.408	C12	1.519	1.071	C8	1.203	2.731				4.109			
		4.071	C26	Valence				4.074			4.125			C13	Valence			
C8	Valence		C25	1.502	1.126			C21	Valence			C2	Valence	C12	1.425	1.414		
C7	1.483	1.191	C27	1.540	1.006	C20	1.408	1.488	C1	1.205	2.715	C1	1.205	2.715	C14	1.199	2.764	
C9	1.514	1.087	C28	1.526	1.049	C22	1.400	1.523	C3	1.424	1.419	C3	1.424	1.419				4.178
C10	1.336	1.841	C29	1.531	1.033	C26	1.520	1.067			4.134			C14	Valence			
		4.119			4.214			4.078	<b>Chen et al. 2009</b>			C13	1.199	2.764				
C8	Valence		C20	Valence				C25	Valence	C2	Valence	C1	1.419	1.440				
C7	1.465	1.256	C19	1.515	1.083	C20	1.406	1.496	C1	1.431	1.390	C1	1.431	1.390			4.204	
C9	1.521	1.064	C21	1.490	1.167	C24	1.392	1.560	C3	1.190	2.839	C3	1.190	2.839	<b>Chiang et al. 2001</b>			
C10	1.347	1.782	C22	1.529	1.039	C29	1.520	1.067			4.228			C1	Valence			
		4.103	C23	1.513	1.090			4.124			C3	Valence	C1a	1.568	0.926			
C8	Valence				4.379			C35	Valence	C2	1.190	2.839	O4	1.254	1.499			
C7	1.476	1.216	C25	Valence		C34	1.408	1.488	C2	1.190	2.839	O3	1.251	1.515				
C9	1.511	1.096	C24	1.338	1.831	C36	1.394	1.551	C4	1.440	1.353				3.940			
C10	1.339	1.825	C26	1.502	1.126	C40	1.517	1.077			4.192	<b>DiPasquale et al. 2006</b>						
		4.137	O1	1.365	1.007			4.115			C6	Valence	C11	Valence				
C8	Valence				3.963			C39	Valence	C5	1.430	1.394	C12	1.522	1.061			
C7	1.481	1.198	C19	Valence		C34	1.408	1.488	C7	1.201	2.747	C7	1.201	2.747	C13	1.520	1.067	
C9	1.468	1.245	C18	1.338	1.831	C38	1.396	1.541			4.141	C6	1.201	2.747	C14	1.525	1.052	
C10	1.352	1.756	C20	1.515	1.083	C43	1.518	1.074	C6	1.201	2.747	C8	1.436	1.369	O2	1.458	0.721	
		4.200	O1	1.357	1.036			4.103	C8	1.436	1.369				3.901			
<b>Andreu et al. 2009</b>					3.950			C4	Valence			4.117	<b>Draskovic et al. 2010</b>					
C4	Valence		C17	Valence				C9	Valence			C2	Valence	C2	Valence			
C3	1.442	1.345	C1	1.391	1.564	C3	1.419	1.440	C8	1.430	1.394	C3	1.436	1.369	C3	1.436	1.369	
O2	1.380	0.954	C18	1.420	1.436	C5	1.397	1.537	C10	1.190	2.839	C7	1.380	1.616	C7	1.380	1.616	
O3	1.215	1.724	C24	1.424	1.419	C9	1.517	1.077			4.232	O1	1.364	1.010	O1	1.364	1.010	
		4.023			4.419			4.054			C10	Valence				3.996		
C21	Valence		C8	Valence				C8	Valence			C9	Valence			C3	Valence	
C16	1.504	1.119	C7	1.498	1.139	C3	1.422	1.427	C8	1.430	1.394	C9	1.190	2.839	C2	1.436	1.369	
C22	1.537	1.015	C9	1.374	1.645	C7	1.392	1.560	C10	1.190	2.839	C11	1.440	1.353	C4	1.429	1.398	
C23	1.536	1.018			4.415	C12	1.512	1.093			4.192				4.278			
C24	1.527	1.045						4.080			C13	Valence	O2	1.287	1.332			
		4.198	<b>Basuli et al. 2003</b>					C21	Valence	C12	1.430	1.394				4.099		
C17	Valence		C20	1.419	1.440	C20	1.419	1.440	C14	1.201	2.747	C14	1.201	2.747			C4	Valence
C14	1.508	1.106	C5	Valence		C22	1.398	1.532			4.141			C3	Valence			
C18	1.516	1.080	C4	1.421	1.431	C29	1.520	1.067			C14	Valence	C3	1.429	1.398			
C19	1.535	1.021	C6	1.391	1.564			4.040			C13	Valence	C5	1.419	1.440			
C20	1.530	1.036	C13	1.518	1.074			C35	Valence	C13	1.201	2.747	C8	1.419	1.440			
		4.243			4.070			4.117	C1	1.436	1.369				4.278			
C14	Valence					C34	1.417	1.448			4.117			C10	Valence			

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C9	1.537	1.015	C3	1.190	2.839	C	Valence			4.409	
O3	1.246	1.543	C5	1.430	1.394	C1	1.376	1.636	C10	Valence	
O4	1.265	1.441			4.232	C	1.386	1.588	C4	1.398	1.532
		3.999		C9	Valence	O	1.378	0.961	C5	1.419	1.440
<b>Gavenonis and Tilley</b>			C8	1.420	1.436			4.184	C9	1.415	1.457
<b>2004</b>			C10	1.140	3.292	C	Valence			4.429	
	C2	Valence			4.728	C4	1.374	1.645		C9	Valence
C1	1.412	1.470		C3	Valence	C	1.386	1.588	C8	1.371	1.660
C3	1.397	1.537	C2	1.192	2.822	O	1.379	0.958	C10	1.415	1.457
C5	1.518	1.074	C4	1.475	1.220			4.191	O1	1.386	0.934
		4.081			4.041	C	Valence			4.051	
	C9	Valence		C11	Valence	C6	1.376	1.636		C2	Valence
C8	1.428	1.402	C10	1.186	2.872	C	1.388	1.578	C3	1.437	1.365
C10	1.405	1.501	C12	1.459	1.279	O	1.377	0.964	O1	1.373	0.978
C14	1.502	1.126			4.151			4.178	O2	1.237	1.593
		4.029	<b>Krawczuk and Stadnicka</b>			C	Valence			3.937	
	C13	Valence	<b>2012</b>			C9	1.379	1.621		C10	Valence
C8	1.404	1.505		C2	Valence	C	1.388	1.578	C4	1.408	1.487
C12	1.399	1.528	C3	1.356	1.738	O	1.378	0.961	C5	1.411	1.474
C17	1.520	1.067	C9	1.490	1.168			4.161	C9	1.405	1.500
		4.101	O1	1.347	1.074	C	Valence			4.461	
	C6	Valence			3.981	C1	1.430	1.394		C9	Valence
C5	1.407	1.492		C5	Valence	C	1.400	1.523	C8	1.392	1.561
C7	1.396	1.541	C4	1.458	1.284	O	1.350	1.062	C10	1.405	1.500
C11	1.522	1.061	C6	1.356	1.737			3.979	O1	1.372	0.980
		4.095	O8	1.351	1.061	C	Valence			4.042	
	C10	Valence			4.081	C4	1.320	1.931		C2	Valence
C5	1.411	1.474	<b>Lee et al. 2006</b>			C	1.400	1.523	C3	1.453	1.303
C9	1.409	1.483		C9	Valence	O	1.400	0.888	O1	1.395	0.906
C14	1.517	1.077	C8	1.522	1.062			4.342	O2	1.211	1.748
		4.035	C13	1.516	1.080	C	Valence			3.957	
	C18	Valence	C14	1.521	1.065	C6	1.400	1.523		C10	Valence
C17	1.421	1.431	O2	1.457	0.725	C	1.380	1.616	C4	1.438	1.361
C19	1.392	1.560			3.933	O	1.340	1.101	C5	1.408	1.488
C23	1.507	1.109	<b>Leon et al. 2002</b>					4.241	C9	1.409	1.483
		4.101		C	Valence	C	Valence			4.332	
	C22	Valence	C1	1.386	1.588	C9	1.420	1.436		C9	Valence
C17	1.416	1.453	C	1.387	1.583	C	1.380	1.616	C8	1.370	1.665
C21	1.380	1.616	O	1.378	0.961	O	1.360	1.025	C10	1.409	1.483
C26	1.530	1.036			4.132			4.077	O1	1.379	0.958
		4.105		C	Valence	<b>Liu et al. 2002</b>				4.106	
	C30	Valence	C4	1.387	1.583		C3	Valence		C2	Valence
C29	1.419	1.440	C	1.387	1.583	C1	1.508	1.106	C3	1.442	1.345
C31	1.393	1.555	O	1.387	0.930	C4	1.211	2.667	O1	1.382	0.947
C35	1.516	1.080			4.097			3.773	O2	1.206	1.781
		4.075		C	Valence		C4	Valence		4.073	
	C34	Valence	C1	1.378	1.626	C3	1.211	2.667		C10	Valence
C29	1.410	1.479	C	1.390	1.569	C5	1.453	1.302	C4	1.435	1.373
C33	1.386	1.588	O	1.369	0.992			3.969	C5	1.408	1.488
C38	1.526	1.049			4.188		C3	Valence	C9	1.398	1.532
		4.115		C	Valence	C2	1.339	1.825		4.393	
<b>Hill et al. 2007</b>			C4	1.369	1.670	C4	1.506	1.113		C9	Valence
	C11	Valence	C	1.390	1.569	C9	1.495	1.149	C8	1.378	1.626
C10	1.383	1.602	O	1.359	1.029			4.087	C10	1.398	1.532
C12	1.226	2.551			4.268		C5	Valence	O1	1.381	0.951
		4.153		C	Valence	C4	1.453	1.302		4.109	
	C12	Valence	C6	1.358	1.725	C6	1.401	1.519		C2	Valence
C11	1.226	2.551	C	1.377	1.631	C10	1.399	1.528	C3	1.448	1.321
C13	1.415	1.457	O	1.391	0.917			4.348	O1	1.375	0.971
		4.008			4.273		C1	Valence	O2	1.211	1.749
	C3	Valence		C	Valence	C2	1.573	0.912		4.042	
C4	1.190	2.839	C9	1.355	1.741	C3	1.508	1.106		C10	Valence
C31	1.420	1.436	C	1.377	1.631	C11	1.529	1.039	C4	1.430	1.394
		4.274	O	1.384	0.941	O1	1.445	0.756	C5	1.406	1.496
	C4	Valence			4.312			3.813	C9	1.401	1.519

**Bond Length – Bond Valence Relationships for C-C and C-O Bonds**

C37	1.377	1.631	C2	1.407	1.492	C7	Valence	O20	1.220	1.694	C8	1.391	1.564							
C41	1.405	1.501	C9	1.415	1.457	C6	1.497	1.143	O21	1.336	1.117		4.232							
		4.180	C10	1.473	1.227	O1	1.314	1.209		4.014		C7	Valence							
	C41	Valence			4.176	O2	1.227	1.652		C2	Valence	C2	1.398	1.532						
C28	1.529	1.039		C6	Valence			4.003	C3	1.458	1.284	C6	1.382	1.607						
C36	1.405	1.501	C7	1.393	1.555		<b>Mattar et al. 2004</b>		O1	1.389	0.923	O	1.407	0.866						
C40	1.372	1.655	C7#1	1.393	1.555	C13	Valence		O11	1.214	1.733		4.005							
		4.195	C6#2	1.512	1.093	C5	1.394	1.550		3.941		C2	Valence							
	C29	Valence			4.203	C9	1.490	1.166		C19	Valence	C1	1.503	1.123						
C28	1.390	1.569		<b>Manbeck et al. 2012</b>		C14	1.392	1.559		C3	1.473	1.228	C3	1.396	1.541					
C30	1.428	1.402		C2	Valence			4.276	O20A	1.223	1.676	C7	1.398	1.532						
C34	1.435	1.373	C1	1.504	1.119		C14	Valence	O21	1.340	1.100		4.196							
		4.344	C3	1.396	1.541		C8	1.396		4.005		C10	Valence							
	C34	Valence	C7	1.395	1.546		C10	1.489				C9	1.365	1.690						
C29	1.435	1.373			4.207		C13	1.392				O1	1.382	0.947						
C33	1.418	1.444		C9	Valence			4.272		<b>Pandey et al. 2014</b>			4.094							
C35	1.422	1.427	C8	1.514	1.087		<b>Mehn et al. 2003</b>			C6	Valence		C11	1.415	1.457					
		4.244	C10	1.389	1.574		C1	Valence		C7	1.491	1.163	O1	1.382	0.947					
	C36	Valence	C14	1.393	1.555		C2	1.569	0.923	O3	1.339	1.105		4.094						
C35	1.453	1.302			4.216		O1	1.278	1.376	O4	1.200	1.820		C2	1.419	1.440				
C37	1.405	1.501		C1	Valence		O3	1.194	1.859			4.088		C6	1.421	1.431				
C41	1.408	1.488		C2	1.393	1.555			4.158	C1	1.520	1.067	C10	1.415	1.457					
		4.290		C6	1.361	1.710		C1	Valence	O1	1.261	1.462		4.328						
	C2	Valence		C7	1.511	1.096		C2	1.493	1.156	O2	1.231	1.628		C6	Valence				
C1	1.521	1.064			4.361		O1	1.285	1.341			4.157		C5	1.411	1.474				
C3	1.414	1.461		C6	Valence		O2	1.240	1.576		C18	Valence	C7	1.423	1.423					
C7	1.366	1.685		C1	1.414	1.461			4.074	C19	1.489	1.170	C11	1.421	1.431					
		4.210		C5	1.391	1.564		C1	Valence	O9	1.341	1.097		4.329						
	C7	Valence		C7	1.497	1.143		C2	1.505	1.116	O10	1.205	1.787		C21	Valence				
C2	1.366	1.685			4.169		O1	1.263	1.452			4.055		C20	1.348	1.777				
C6	1.374	1.645		C3	Valence		O2	1.237	1.593		C17	Valence	C22	1.421	1.431					
C8	1.539	1.009		C4	1.530	1.036			4.161		O7	1.264	1.446	O2	1.404	0.875				
		4.339		O3	1.281	1.361		C1	Valence		O8	1.237	1.593		4.084					
	C14	Valence		O4	1.227	1.652		C2	1.528	1.042			4.095		C22	Valence				
C1	1.530	1.036			4.049		O1	1.201	1.813		<b>Reddy et al. 2014</b>		C13	1.524	1.055		4.084			
C9	1.382	1.607		C1	Valence		O2	1.251	1.515		C1	Valence	O7	1.264	1.446		C22	Valence		
C13	1.390	1.569		C2	1.517	1.077			4.371		C2	1.487	1.177		4.095		C13	1.426	1.410	
		4.212		O1	1.261	1.462		<b>Munshi et al. 2010</b>			O1	1.297	1.285				C17	Valence		
	C21	Valence		O2	1.237	1.593		C3	Valence		O2	1.203	1.800				C16	1.406	1.496	
C16	1.317	1.948			4.132		C2	1.454	1.299			4.262		C18	1.425	1.414		4.385		
C20	1.425	1.414		C8	Valence		C4	1.379	1.620		<b>Saeed et al. 2012</b>			4.385			C10	1.361	1.710	
C22	1.552	0.971		C9	1.514	1.087		C19	1.480	1.204		C11	Valence				C12	1.425	1.414	
		4.333		O3	1.291	1.313			4.122		C1	1.455	1.293				O1	1.391	0.917	
	C16	Valence		O4	1.225	1.663			C8	Valence		C15	1.355	1.742				4.052		
C15	1.623	0.787			4.063		C7	1.437	1.364		O12	1.362	1.017					4.042		
C17	1.371	1.660		C1	Valence		C9	1.505	1.116				4.052					C31	Valence	
C21	1.317	1.948		C2	1.504	1.119		C18	1.423	1.423			3.903				C2	1.459	1.280	
		4.395		O1	1.278	1.376				3.903		C2	1.459	1.280				C35	1.357	1.730
	C23	Valence		O2	1.242	1.565			C3	Valence		O32	1.371	0.985				C32	1.371	0.985
C22	1.543	0.997			4.060		C2	1.458	1.284				3.994						4.299	
C24	1.346	1.788		C1	Valence		C4	1.382	1.607		<b>Schumann et al. 2003</b>									
C28	1.401	1.519		C2	1.501	1.129		C19	1.473	1.228		C14	Valence							
		4.304		O1	1.291	1.313				4.119		C9	1.376	1.636						
	C28	Valence		O2	1.228	1.646			C8	Valence		C13	1.381	1.612						
C15	1.524	1.055			4.088		C7	1.438	1.361			O	1.435	0.783						
C23	1.401	1.519		C3	Valence		C9	1.506	1.112				4.030							
C27	1.367	1.680		C4	1.506	1.113		C18	1.419	1.438				4.030						
		4.253		O1	1.336	1.117				3.911			C8	Valence						
	<b>Maher et al. 2010</b>			O2	1.203	1.800			C2	Valence		C3	1.391	1.564						
	C10	Valence			4.030				C3	1.454	1.299		C7	1.378	1.626					
C1	1.473	1.227			C7	Valence			O1	1.392	0.915		O	1.404	0.875					
O1	1.348	1.070			C1	1.511	1.096		O11	1.211	1.747			4.066						
O2	1.211	1.749			O1	1.280	1.366				3.960			C3	Valence					
		4.046			O2	1.221	1.688							C2	1.502	1.126				
	C1	Valence				4.149			C3	1.480	1.204			C4	1.396	1.541				



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## Literature Cited

- Alonso M, R Chicharro, C Miranda, VJ Arán, MA Maestro, and B Herradón.** 2009. X-ray Diffraction, Solution Structure, and Computational Studies on Derivatives of (3-*sec*-Butyl-2,3-dihydro-1*H*-isoquinolin-4-ylidene)acetic Acid: Compounds with Activity as Calpain Inhibitors. *Journal of Organic Chemistry* 75:342-52.
- Allen FH, O Kennard, DG Watson, L Brammer, AG Orpen, and R Taylor.** 1987. Tables of Bond Lengths determined by X-Ray and Neutron Diffraction. Part 1. Bond Lengths in Organic Compounds. *Journal of the Chemical Society Perkin Transactions II* S1-S19.
- Andreu R, L Carrasquer, S Franco, J Garín, J Orduna, NM de Baroja, R Alicante, et al.** 2009. 4*H*-Pyran-4-ylidenes: Strong Proaromatic Donors for Organic Nonlinear Optical Chromophores. *Journal of Organic Chemistry* 74:6647-57.
- Basuli F, J Tomaszewski, JC Huffman, and DJ Mindiola.** 2003. Carbon-Oxygen Bond Cleavage Promoted by a Scandium Borohydride Complex. *Organometallics* 22: 4705-14.
- Bleeke JR, BL Lutes, M Lipschutz, D Sakellariou-Thompson, and JS Lee.** 2010 Synthesis, Structure, Spectroscopy, and Reactivity of Oxapentadienyl-Cobalt-Phosphine Complexes. *Organometallics* 29:5057-67.
- Bleeke JR and W Anutrasakda.** 2012. Synthesis, Structure, Spectroscopy, and Reactivity of Azapentadienyl-Cobalt-Phosphine Complexes. *Organometallics* 31:2219-30.
- Borbulevych OY, OV Shishkin, and MY Antipin.** 2002. X-ray Diffraction and Ab Initio Quantum-Chemical Study of the Charge Density in the Crystals of Meisenheimer Complexes – Derivatives of 2,4,6-Trinitrobenzene and 5,7-Dinitroquinoline. *Journal of Physical Chemistry* 106:8109-16.
- Breese NE and M O’Keeffe.** 1991. Bond-Valence Parameters for Solids. *Acta Crystallographica* B47:192-7.
- Brown ID.** 2002. *The Chemical Bond in Inorganic Chemistry.* Oxford University Press (New York). Chapter 9, Physical properties of bonds: bond lengths and bond angles; 106-9.
- Brown ID and D Altermatt.** 1985. Bond-valence parameters obtained from a systematic analysis of the inorganic crystal structure database. *Acta Crystallographica.* B41: 244-247.
- Burlakov VV, VS Bogdanov, P Arndt, W Baumann, A Spannenberg, KA Lyssenko, IV Ananyev, et al.** 2015. Thermal Isomerization of the Buchwald Seven-Membered Zirconacyclocumulene and Its Interaction with Acetylenes. Synthesis and Structures of Novel Seven-Membered Zirconacyclocumulene Complexes. *Organometallics* 34:2471-80.
- Chen G, L Dawe, L Wang, and Y Zhao.** 2009. Planar Acetylene-Expanded TTFAQ Analogues. *Organic Letters* 11:2736-9.
- Chiang RK, CC Huang, and CS Wur.** 2001. Hydrothermal Synthesis and Structural Characterization of a Coordination Polymer Containing Heptanuclear Co<sub>7</sub>(μ<sub>3</sub>-OH)<sub>8</sub> Clusters. *Inorganic Chemistry* 40:3237-9.
- Clementi E and DL Raimondi.** 1963. Atomic screening constants from SCF functions. *Journal of Chemical Physics* 38:2686-9.
- Clementi E, DL Raimondi, and WP Reinhardt.** 1967. Atomic screening constants from SCF functions. II Atoms with 37 to 86 electrons. *Journal of Chemical Physics* 47:1300-7.
- DiPasquale AG, DA Hrovat, and JM Mayer.** 2006. Non-Redox-Assisted Oxygen-Oxygen Bond Homolysis in Titanocene Alkylperoxide Complexes: [Cp<sub>2</sub>Ti<sup>IV</sup>(η<sup>1</sup>-OO<sup>t</sup>Bu)L]<sup>+0</sup>, L = Cl, OTf<sup>-</sup>, Br<sup>-</sup>, OEt<sub>2</sub>, Et<sub>3</sub>P. *Organometallics* 25:915-24.
- Drašковиć BM, GA Bogdanović, MA Neelakantan, AC Chamayou, S Thalamuthu, YS Avadhut, JS auf der Günne, et al.** 2010. N-*o*-Vanillylidene-L-histidine: Experimental Charge Density Analysis of a Double Zwitterionic Amino Acid Schiff-Base Compound. *Crystal Growth & Design* 10:1665-76.
- Esteruelas MA, C García-Yebra, M Oliván, E Oñate, and MA Tajada.** 2000. The Dihydride-Osmium(IV) Complex [OsH<sub>2</sub>(κ<sup>2</sup>-O<sub>2</sub>CCH<sub>3</sub>)(H<sub>2</sub>O)(P<sup>i</sup>Pr<sub>3</sub>)<sub>2</sub>]BF<sub>4</sub> as a Precursor for Carbon-Carbon Coupling Reactions. *Organometallics* 19:5098-106.
- Gavenonis J and TD Tilley.** 2004. Synthesis and Reactivity of Tris(imido)rhenium Complexes Containing Rhenium-Main Group Element Bonds. Silicon-Carbon Bond Activations of PhSiH<sub>3</sub> by Silyl Complexes. *Inorganic Chemistry* 43:4353-62.
- Hardcastle FD.** 2013. Valence-Length Correlations for Chemical Bonds from Atomic Orbital Exponents. *Journal of the Arkansas Academy of Science* 67:53-8.



- Hill AF, AD Rae, M Schultz, and AC Willis.** 2007. Bis(alkynyl), Metallacyclopentadiene, and Diphenylbutadiyne Complexes of Ruthenium. *Organometallics* 26:1325-38.
- Huynh L, Z Wang, J Yang, V Stoeva, A Lough, I Manners, and MA Winnik.** 2005. Evaluation of Phosphorescent Rhenium and Iridium Complexes in Polythionylphosphazene Films for Oxygen Sensor Applications. *Chemistry of Materials* 17:4765-73.
- Krawczuk A and K Stadnicka.** 2012. Experimental and Theoretical Charge Density Study of the Chemical Bonding in Chlorokojic Acid Crystal Structure. *Journal of Physical Chemistry A* 116:9759-68.
- Lee SB, AC Willis, and RD Webster.** 2006. Synthesis of the Phenoxonium Cation of an  $\alpha$ -Tocopherol Model Compound Crystallized with Non-Nucleophilic  $[B(C_6F_5)_4]^-$  and  $(CB_{11}H_6Br_6)^-$  Anions. *Journal of the American Chemical Society* 128:9332-3.
- Leon LA, R Notario, J Quijano, and C Sánchez.** 2002. Structures and Enthalpies of Formation in the Gas Phase of the Most Toxic Polychlorinated Dibenzo-*p*-dioxins. A DFT Study. *Journal of Physical Chemistry A* 106:6618-27.
- Liu JF, SL Huang, YC Lin, YH Liu, and Y Wang.** 2002. New Acetylde Migration and Oxygen Transfer Reactions in Ruthenium Complexes Containing an Acetyl-Substituted Cp Ligand. *Organometallics* 21:1355-61.
- Liu X, JM Cole, PG Waddell, TC Lin, and S McKechnie.** 2013. Molecular Origins of Optoelectronic Properties in Coumarins 343, 314T, 445, and 522B. *Journal of Physical Chemistry C* 117:14130-41.
- Liu X, Z Su, W Ji, S Chen, Q Wei, G Xie, X Yang, and S Gao.** 2014. Structure, Physicochemical Properties, and Density Functional Theory Calculation of High-Energy-Density Materials Constructed with Intermolecular Interaction: Nitro Group Charge Determines Sensitivity. *Journal of Physical Chemistry C* 118:23487-98.
- Lu CC and JC Peters.** 2006. Pseudotetrahedral Manganese Complexes Supported by the Anionic Tris(phosphino)borate Ligand  $[PhBP^{iPr}_3]$ . *Inorganic Chemistry* 45:8597-607.
- Maher TR, AD Spaeth, BM Neal, CL Berrie, WH Thompson, VW Day, and MV Barybin.** 2010. Linear 6,6'-Biazulenyl Framework Featuring Isocyanide Termini: Synthesis, Structure, Redox Behavior, Complexation, and Self-Assembly on Au(111). *Journal of the American Chemical Society* 132:15924-6.
- Manbeck KA, S Kundu, AP Walsh, WW Brennessel, and WD Jones.** 2012. Carbon-Oxygen Bond Activation in Esters by Platinum(0): Cleavage of the Less Reactive Bond. *Organometallics* 31:5018-24.
- Mattar SM, AH Emwas, and LA Calhoun.** 2014. Spectroscopic Studies of the Intermediates in the Conversion of 1,4,11,12-Tetrahydro-9,10-anthraquinone to 9,10-Anthraquinone by Reaction with Oxygen under Basic Conditions. *Journal of Physical Chemistry A* 108:11545-53.
- Mehn MP, K Fujisawa, EL Hegg, and L Que, Jr.** 2003. Oxygen Activation by Nonheme Iron(II) Complexes:  $\alpha$ -Keto Carboxylate versus Carboxylate. *Journal of the American Chemical Society* 125:7828-42.
- Munshi P, C Jelsch, VR Hathwar, and TNG Row.** 2010. Experimental and Theoretical Charge Density Analysis of Polymorphic Structures: The Case of Coumarin 314 Dye. *Crystal Growth & Design* 10:1516-26.
- Pandey AK, GPA Yap, and NJ Zondlo.** 2014. (2*S*,4*R*)-4-Hydroxyproline(4-nitrobenzoate): Strong Induction of Stereoelectronic Effects via a Readily Synthesized Proline Derivative. Crystallographic Observation of a Correlation between Torsion Angle and Bond Length in a Hyperconjugative Interaction. *Journal of Organic Chemistry* 79:4174-9.
- Pandian RP, NP Raju, JC Gallucci, PM Woodward, AJ Epstein, and P Kuppasamy.** 2010. A New Tetragonal Crystalline Polymorph of Lithium Octa-*n*-Butoxy-Naphthalocyanine (LiNc-BuO) Radical: Structural, Magnetic and Oxygen-Sensing Properties. *Chemistry of Materials* 22:6254-62.
- Pauling L.** 1929. The principles determining the structure of complex ionic crystals. *Journal of the American Chemical Society* 51:1010-26.
- Pauling L.** 1947. Atomic radii and interatomic distances in metals. *Journal of the American Chemical Society* 69:542-53.

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- Pierce GA, D Vidovic, DL Kays, ND Coombs, AL Thompson, ED Jemmis, S De, and S Aldridge.** 2009. Half-Sandwich Group 8 Borylene Complexes: Synthetic and Structural Studies and Oxygen Atom Abstraction Chemistry. *Organometallics* 28:2947-60.
- Reddy ST, KP Krovi, and MJ Swamy.** 2014. Structure and Thermotropic Phase Behavior of a Homologous Series of *N*-Acylglycines: Neuroactive and Antinociceptive Constituents of Biomembranes. *Crystal Growth & Design* 14:4944-54.
- Saeed A, MF Erben, and M Bolte.** 2012. Twisted Imide Bond in Noncyclic Imides. Synthesis and Structural and Vibrational Properties of *N,N*-Bis(furan-2-carbonyl)-4-chloroaniline. *Journal of Organic Chemistry* 77:4688-95.
- Schomaker VS and DP Stevenson.** 1941. Some Revisions of the Covalent Radii and the Additivity Rule for the Lengths of Partially Ionic Single Covalent Bonds. *Journal of the American Chemical Society* 63:37-40.
- Schumann H, S Dechert, S Schutte, JY Hyeon, M Hummert, and BC Wassermann.** 2003. Oxygen-Stabilized Organoaluminum Compounds as Highly Active Cocatalysts for Ziegler-Natta Olefin Polymerization. *Organometallics* 22:1391-401.
- Song J, YX Lei, and Z Rappoport.** 2007. The First Solid Enols of Anhydrides. Structure, Properties, and Enol/Anhydride Equilibria. *Journal of Organic Chemistry* 72:9152-62.
- Varga V, I Císařová, R Gyepes, M Horáček, J Kubišta, and K Mach.** 2009. Evaluation of the Oxygen  $\pi$ -Donation in Permethylytitanocene Silanolates and Alcoholates. *Organometallics* 28:1748-57