

## Reaction of SO<sub>2</sub> and SO<sub>2</sub> with O<sub>2</sub> after dissolution during benchtop experiments of CO<sub>2</sub> storage at elevated temperature and pressure

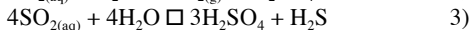
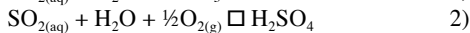
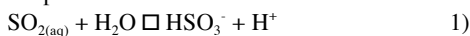
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Captured CO<sub>2</sub> streams from the burning of fossil fuels can contain impurities like SO<sub>2</sub> and O<sub>2</sub> as well as others. The presence of these impurities in CO<sub>2</sub> that is being geologically stored can result in significant impacts on the chemistry of the aqueous phase. Understanding how these changes in chemistry may affect the physical and chemical behaviour of the storage system is a vital component of the risk assessment process. Fundamental to that task is being able to correctly predict how the impurities react once dissolved in the water because the formation of strong acids and changes in redox state are expected to occur. When SO<sub>2</sub> dissolves in water there are several possible reactions that can occur:



1) is a simple hydration reaction that is reversible and can result in low pH. The second reaction involves oxidation (by O<sub>2</sub> or other oxidants) and results in lower pH than 1). 3) is a disproportionation reaction that will produce both oxidized and reduced sulphur species and will result in pH higher than 2) but lower than 1). The H<sub>2</sub>S in 3) can be substituted for by elemental S as a product. A series of batch experiments were conducted to evaluate the effects of impurities on CO<sub>2</sub> storage in sedimentary rocks from the Surat Basin, Australia (Farquhar et al., 2015; Pearce et al., 2015). The experiments were run using pure CO<sub>2</sub>, CO<sub>2</sub> + 0.16% SO<sub>2</sub> and CO<sub>2</sub> + 0.16% SO<sub>2</sub> + 2% O<sub>2</sub> at 12 MPa and 60°C. Here we report the observations regarding how SO<sub>2</sub> reacted during the experiments. With 0.16% SO<sub>2</sub> and no additional oxidant, disproportionation did occur with either S<sub>0</sub> or H<sub>2</sub>S or both and very low pH. In the presence of O<sub>2</sub>, the SO<sub>2</sub> was oxidized giving sulphuric acid and typically a lower pH than the experiments without O<sub>2</sub>. During the 15 days of the experiment runs, it was clear there was a rate control on the SO<sub>2</sub> oxidation or disproportionation reactions.

Farquhar et al. (2015) Chem. Geol. 399, 98-122;  
Pearce et al. (2015) Chem. Geol. 399, 65-86.