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Reaction Kinetics, Mechanisms and Catalysis, Volume 121 (2017) 141-160.

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ISSN: 1878-5190; Source Type: Journal; Original language: English; Document Type: Article; Publisher: Springer

Akadémiai Kiadó, Budapest, Hungary 2017

The final publication is available at Springer via
<http://dx.doi.org/10.1007/s11144-017-1155-5>

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Received: 30 November 2016/Accepted: 4 February 2017/Published online: 15 February 2017

Electronic supplementary material:

The online version of this article (doi: 10.1007/s11144-017-1155-5) contains supplementary material, which is available to authorized users.

Effect of Mo incorporation on electrocatalytic performance of Ti-Mo mixed oxide-carbon composite supported Pt electrocatalysts

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Abstract

Ti_(1-x)Mo_xO₂-C carbon composite supported platinum electrocatalysts with systematically varied Ti/Mo ratio were prepared by a multistep sol-gel-based synthesis method. The effect of the composition of the support material on the electrochemical behaviour of the 20 wt.% Pt electrocatalysts was investigated and the samples were also characterised by XRD, TEM, EDX and XPS techniques. The composite support ensures enhanced CO tolerance compared to the reference commercial 20 wt.% Pt/C (Quintech) catalyst. Different Mo species were identified to have critical importance on the electrocatalytic performance both in hydrogen oxidation and CO tolerance. The unincorporated Mo species are not stable upon applying a wide cyclic polarization window and as a consequence they are removed gradually. Higher stability was obtained over Mo species incorporated into the rutile lattice. From the results Ti/Mo= 80/20 atomic ratio has been suggested as an optimal composition having the largest ratio of incorporated/non-incorporated Mo species.

Keywords Conducting Ti-based mixed oxides, TiMoO_x, Composite materials, Pt electrocatalysts, CO-tolerance

Introduction

Fuel cells are well positioned to provide an economically compelling solution for “green” electric power generation. For the wide-range implementation of polymer electrolyte membrane fuel cells (PEMFC), it is most essential to produce low-cost durable units. For fuel cells, state of the art systems contain large quantities of platinum supported on carbon black. However, the electrochemical corrosion of carbon [1,2] and concomitant sintering of Pt [3-5] leads to stability problems jeopardizing long term operation of PEMFC stacks. It is therefore important to explore alternative materials that are more stable compared to carbon, i.e. highly corrosion resistant across the anticipated potential/pH window experienced in fuel cell electrodes. Titanium dioxide is promising candidate as catalyst support in PEMFC because TiO₂ has very good chemical stability in acidic and oxidative environments [6,7]. However, undoped TiO₂ oxide is an n-type semiconductor with low electron conductivity and doping with transition metals (e.g., W, Nb and Ta) has been used to enhance its electrical conductivity [8-11].

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On the other hand, CO tolerance of the Pt-based catalysts is of high importance when feeding low temperature PEMFC stacks with hydrogen from reformates. It has been demonstrated that by adding oxophilic doping metals such as W and Mo the CO tolerance of the Pt/C electrocatalysts is significantly improved. The second metal is reported to (i) provide the necessary OH_{ad} species at less positive potentials than Pt (according to the “bifunctional mechanism”); (ii) modify the electronic structure of Pt and therefore reduce the adsorption strength of CO (the “electronic effect”); and (iii) to increase catalytic activity as the result of the “hydrogen spillover” effect from Pt to tungsten or molybdenum oxides, leading to formation of bronzes (H_xMO_3 , M= W, Mo) [12-16]. It is well-known [7,17] that MO_3 oxides (M= W, Mo) can be electrochemically reduced to a non-stoichiometric and electroconductive H_xMO_3 .

In agreement with the bifunctional mechanism, the co-catalytic activity is supposed to be due to the good redox properties of the oxides involved. Usually, molybdenum oxides do not exist in a single phase but in mixed-valence oxides (MoO_x) [18]. For the Pt-Mo/C system Santiago *et al.* [19] proposed that the increased CO tolerance is a result of synergism between water gas shift reaction and the bifunctional mechanism.

Recently, the combination of the highly stable TiO_2 with MoO_x , was reported to have high electronic conductivity and relative stability in acid solutions. Accordingly, the multifunctional $\text{Ti}_{0.7}\text{Mo}_{0.3}\text{O}_2$ support material was successfully applied for Pt-based catalysts for the oxygen reduction reaction (ORR) [20,21].

In our previous studies [22,23] platinum was deposited onto a $\text{Ti}_{0.7}\text{W}_{0.3}\text{O}_2$ -C composite support in order to stabilize the Pt nanoparticles and increase the CO-tolerance. Optimum experimental conditions were found to synthesize rutile phase, tungsten-doped $\text{Ti}_{0.7}\text{W}_{0.3}\text{O}_2$ -C composites by a novel, multistep sol-gel method resulting in exclusive incorporation of W into TiO_2 -rutile lattice. Better performance in the CO electrooxidation for this catalyst compared to the PtRu/C benchmark was demonstrated. The electrochemical stability tests revealed that the degradation rate of the 40 wt.% Pt/ $\text{Ti}_{0.7}\text{W}_{0.3}\text{O}_2$ -C catalyst is much smaller than that of the commercial 40 wt.% Pt/C (Quintech) catalyst.

It is known that less stable transition metal oxides tend to dissolve under acidic conditions and high potentials. Such leaching and subsequent uncontrolled re-deposition of certain components of the electrocatalyst may be detrimental for the performance of the fuel cell, so stability against initial leaching should be checked.

Our aim in this work was to develop a synthesis method for Mo-containing $\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2$ -C ($x= 0.2-0.4$) composite supports loaded with 20 wt.% Pt. Using a combination of structural, spectroscopic and electrochemical methods we assessed the incorporation limit of Mo into the rutile lattice by the proposed synthesis method. The effect of the incorporated and excess molybdenum on the electrocatalytic properties and the short-term stability is also assessed.

Experimental

Preparation of the $\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2$ -C support materials

Titanium-isopropoxide ($\text{Ti}(\text{O}-i\text{-Pr})_4$, Aldrich, 97%) and ammonium heptamolybdate tetrahydrate ($(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \times 4\text{H}_2\text{O}$, Merck, 99%) were used as Ti and Mo precursor compounds.

The composite support materials consisting of electroconductive $\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2$ mixed oxide and activated carbon were prepared by using a multistep sol-gel-based synthesis method followed by high-temperature heat treatment, described in detail and demonstrated schematically in Fig. S1 of the Supplementary Material.

Table 1 reports data on the actual composition and preparation details of the $\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2\text{-C}$ ($x= 0.2\text{-}0.4$) composites. The sample identifier contains the intended composition of the $\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2$ mixed oxide reflecting the desired Ti/Mo atomic ratio; in all cases the mass ratio of the mixed oxide to the activated carbon was 75:25 (denoted as $\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2\text{-C}$).

Table 1. Nominal composition and preparation details of the samples with the different Ti to Mo ratios.

Samples ^{a)}	Ti/Mo (at/at)	Ti(O- <i>i-Pr</i>) ₄ , ml	1.44 M HNO ₃ , ml	(NH ₄) ₆ Mo ₇ O ₂₄ 4 H ₂ O, mg
Ti _{0.8} Mo _{0.2} O ₂ -C	0.8/0.2	2.1	23.4	299
Ti _{0.7} Mo _{0.3} O ₂ -C	0.7/0.3	1.7	19.4	426
Ti _{0.6} Mo _{0.4} O ₂ -C	0.6/0.4	1.4	15.8	540

^{a)} Expected composition of $\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2$ mixed oxide reflects desired Ti/Mo atomic ratio.

Preparation of the 20 wt.% Pt/Ti_(1-x)Mo_xO₂-C electrocatalysts

$\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2\text{-C}$ composite materials were loaded with 20 wt.% Pt via a modified NaBH₄-assisted ethylene-glycol (EG) reduction-precipitation method (see Fig. S2 and its description in the Supplementary Material) [22-24].

Physical characterization

The powder X-ray diffraction (XRD) patterns were obtained in a Philips model PW 3710 based PW 1050 Bragg-Brentano para-focusing goniometer using CuK_α radiation ($\lambda= 0.15418$ nm), graphite monochromator and proportional counter. For quantitative analysis silicon powder (NIST SRM 640) was used as internal standard and the scans were evaluated with profile fitting methods. The cell parameters of the crystalline phases were determined from the fitted d-values.

Transmission Electron Microscopy (TEM) studies of the samples were made by use of a FEI Morgagni 268D type transmission electron microscope (accelerating voltage: 100 kV, W-filament). The samples were prepared by grinding in an Achate mortar and dispersing of the resulted powder in ethanol. The average diameter and particles size distribution were calculated by measuring the diameters of no less than 800 randomly selected metal particles in at least ten micrographs of each sample taken from non-aggregated areas.

Energy dispersive X-ray spectrometry (EDX) analysis was carried out with an INCA (Oxford Instruments Ltd.) detector and the INCA Energy software package. EDX analysis of samples was performed by using ZEISS EVO 40XVP scanning electron microscope (accelerating voltage: 20 kV, W-filament). At least 5 spatially well separated points have been analyzed in order to get the reported composition values.

X-ray photoelectron spectroscopy (XPS) measurements were carried out using an EA125 electron spectrometer manufactured by OMICRON Nanotechnology GmbH (Germany). The photoelectrons were excited by unmonochromatized MgK_α (1253.6 eV) radiation. Spectra were recorded in the Constant Analyser Energy mode of the energy analyser with 30 eV pass energy resulting in a spectral resolution around 1 eV. The powder catalyst samples were suspended in isopropanol then some of this suspension was dried to a standard OMICRON sample plate. Binding energies were referenced to the main component of the C 1s spectrum of the support (graphite at 284.4 eV binding energy). Data were processed using the CasaXPS software package [25]. Nominal surface compositions were calculated using the XPS MultiQuant software package

[26,27]. Chemical states were identified by XPS databases [28,29] and with the help of the related literature.

Electrochemical characterization

The electrocatalysts were investigated by means of cyclic voltammetry and CO_{ads} stripping voltammetry techniques in a standard three-electrode electrochemical cell. The applied electrolyte was 0.5 M H₂SO₄. The working electrode was prepared by dipping a drop of catalyst ink on a freshly polished glassy carbon electrode (d = 0.3 cm, geometric surface area A = 0.0707 cm²), and air-drying at room temperature for 30 min. The Pt loading of the electrodes was 10 μg cm⁻². For the catalyst ink 5 mg catalyst sample was suspended in 4 ml H₂O + 1 ml isopropanol + 20 μl Nafion solution (D520 Nafion Dispersion - Alcohol based 1000 EW at 5 wt%, DuPont™ Nafion®). The reference electrode was a hydrogen electrode immersed in the same electrolyte as the working electrode and all potentials are given on the RHE scale. Pt wire was used as the counter electrode. In the so-called "pre-leaching" procedure, the samples were submitted to cyclic polarization with 100 mV s⁻¹ for 30 cycles between 50 and 1000 mV potential limits.

In order to get information on the initial behavior of the electrocatalysts, CO_{ads} stripping voltammetry measurements were carried out on the samples without any electrochemical pre-conditioning or cleaning, and after the "pre-leaching" treatment.

For CO_{ads} stripping voltammetry gaseous CO was fed into the cell for 20 min while maintaining a constant potential of 50 mV. Two types of CO_{ads} stripping methods are distinguished according to the way how the physically adsorbed CO from the surface and dissolved CO from the electrolyte was removed. In case of so called Ar-purged and H₂-purged CO_{ads} stripping methods Ar or H₂ streams were bubbled through the electrolyte, respectively. After Ar or H₂ purge for 20 min, the working electrode was subjected to a cyclic voltammetry scans at a 10 mV s⁻¹ scan rate between 50 and 1000 mV. During the first cycle, the chemically adsorbed CO was oxidized (1st CO stripping), then after the CO electrooxidation, another cyclic voltammogram was recorded. After the first CO stripping measurement the electrode was subjected to 30 cycles of the "pre-leaching" procedure at a 100 mV s⁻¹ scan rate between 50 and 1000 mV. Then the above described CO stripping measurement was repeated (2nd CO stripping). After the 2nd CO_{ads} stripping measurement the cyclic voltammogram was recorded again. Relative errors were calculated as the standard deviation of at least three independent measurements.

For comparison as reference commercially available 20 wt.% Pt (Quintech, C-20-Pt, on Vulcan; denoted hereafter as 20Pt/C) and PtRu/C, considered as the state-of-the art CO-tolerant electrocatalyst (Quintech, C-20-/10-Pt/Ru, Pt= 20 wt%, Ru= 10 wt% on Vulcan; denoted hereafter as PtRu/C) were also studied by the same method as described above.

Results and discussion

Characterization of the Ti_(1-x)Mo_xO₂-C composite materials and the related Pt catalysts

The structure of the Ti_(1-x)Mo_xO₂-C composite materials was investigated by XRD measurements to evaluate the phase composition of the materials and determine the lattice parameters. Fig. 1 shows the XRD patterns of the composite samples before (Fig. 1A) and after the high-temperature treatment (HTT) (Fig. 1B). The results are also summarized in Table 2.

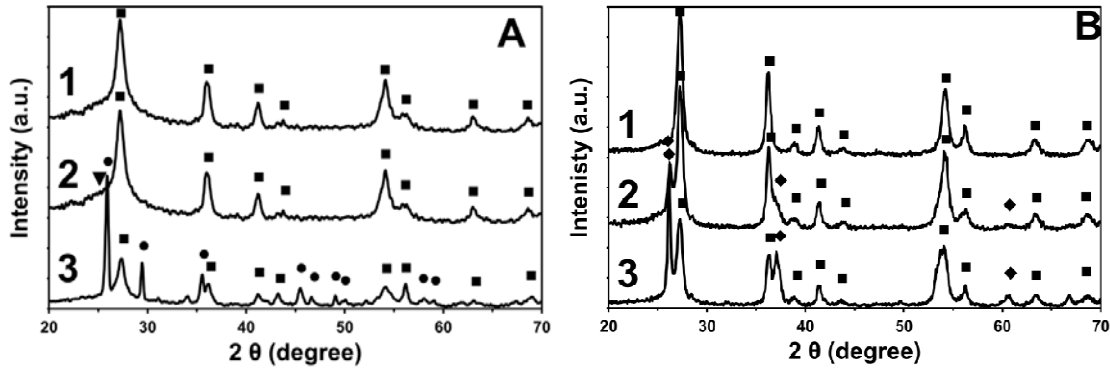


Fig. 1 XRD patterns of $Ti_{(1-x)}Mo_xO_2-C$ samples before (A) and after high-temperature treatment (B). 1: $Ti_{0.8}Mo_{0.2}O_2-C$; 2: $Ti_{0.7}Mo_{0.3}O_2-C$; 3: $Ti_{0.6}Mo_{0.4}O_2-C$. ■ - rutile; ▼ - anatase; ● - MoO_3 ; ◆ - MoO_2 .

Upon preparation of the $Ti_{0.7}W_{0.3}O_2-C$ composite materials we have found [22,23] that preliminary formation of the rutile- TiO_2 phase is a necessary condition for complete isovalent W incorporation into the rutile lattice upon the following reductive treatment. Accordingly, in the course of preparation of Mo-containing composite materials, our goals were: (i) to produce $Ti_{(1-x)}Mo_xO_2$ rutile phase with high crystallinity and (ii) to transform of Mo into oxidation state of four. Computer modelling of a $Ti_{(1-x)}W_xO_2$ mixed oxide revealed [30] that the unit cell parameters and volume of the rutile lattice change significantly, as a result of tungsten incorporation. Therefore, it is assumed that the rutile lattice distortion can be used as an indicator of the dopant incorporation.

The samples obtained before HTT consist of rutile crystallites (Fig. 1A). In the case of sample $Ti_{0.6}Mo_{0.4}O_2-C$ a small amount of MoO_3 was also detected (see Table 2), due to the greater amount of molybdenum precursor used upon the preparation (rutile/ MoO_3 = 72/28). At the other two samples with less nominal Mo content there were no reflections characteristic to MoO_3 found, which may be attributed to the presence of amorphous MoO_3 or crystalline MoO_3 below the detection limit.

Table 2. Structural properties of the $Ti_{(1-x)}Mo_xO_2-C$ composite materials determined by XRD analysis.

Samples ^{a)}	WHTT ^{b)} (Phase, %)			HTT ^{c)} (Phase, %)			Lattice parameters, Å ^{d)}	Mo subst., %
	R	A	MoO_3/MoO_2	R	A	MoO_3/MoO_2		
$Ti_{0.8}Mo_{0.2}O_2-C$	100	0	0/0	90	10	0/0	$a=4.640, c= 2.935$	21
$Ti_{0.7}Mo_{0.3}O_2-C$	90	10	0/0	86	10	0/4	$a=4.645, c= 2.932$	25
$Ti_{0.6}Mo_{0.4}O_2-C$	72	0	28/0	80	5	0/15	$a=4.655, c= 2.928$	30

^{a)} Expected composition of $Ti_{(1-x)}Mo_xO_2$ mixed oxide reflects desired Ti/Mo atomic ratio;

^{b)} WHTT (without high-temperature treatment): the samples were studied before any HTT;

^{c)} Detailed description of the HTT applied was shown in Fig. 1 (see Experimental part);

^{d)} Lattice parameters of the rutile phase obtained after HTT; Pure rutile TiO_2 : $a= 4.593 \text{ \AA}, c= 2.959 \text{ \AA}$;

R: rutile, A: anatase.

After the heat treatment (see Fig. 1B) almost pure rutile phase was obtained along with a small amount of both the anatase (5-10 %) and MoO₂ (4-15 %) phases. The lattice parameters of the rutile phase after heat treatment are also given in Table 2. A comparison with the lattice parameters of pure rutile TiO₂ indicates the expected distortion and confirm the presence of molybdenum incorporated into the unit cell; the degree of Mo incorporation increases to quite high values (Mo_{subst}= 21-30 %). The presence of the MoO₂ phase indicates, however, that a certain fraction of Mo ions remains outside of the titania lattice and condensates into Mo-oxide particles located on the surface of the composite support. According to the XRD data, the amount of free, i.e. unincorporated MoO₂ phase also increased with increasing nominal molybdenum content. It seems that incorporation of molybdenum under the used conditions has a maximum.

Transmission electron microscopy (TEM) investigations

TEM micrograph, electron diffraction pattern and particle size distribution of the 20 wt.% Pt/Ti_{0.8}Mo_{0.2}O₂-C catalyst are shown in Figure 2. Very similar morphologies were observed for the Pt/Ti_{0.7}Mo_{0.3}O₂-C and the Pt/Ti_{0.6}Mo_{0.4}O₂-C catalysts; micrographs, electron diffraction patterns and histograms for particle size distribution for these samples are presented in Figs. S3 and S4 of the Supplementary Material.

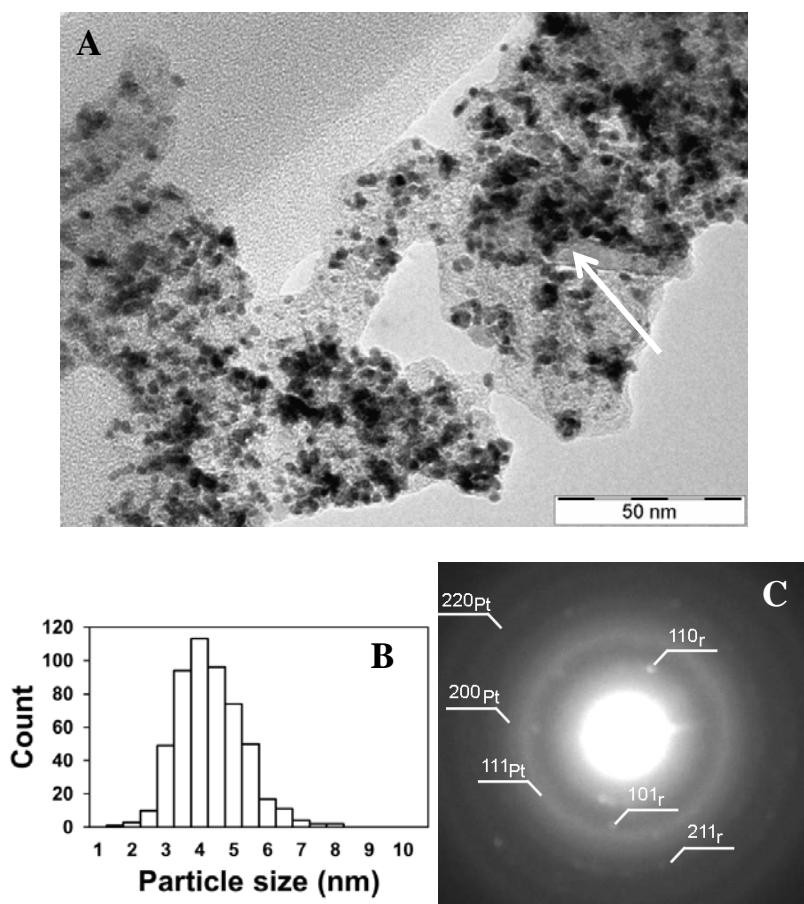


Fig. 2 TEM images (A), histogram of the Pt particle size distribution (B) and electron diffraction pattern (C) of the Pt/Ti_{0.8}Mo_{0.2}O₂-C electrocatalyst; “r” stands for the Mo-doped rutile phase.

Figs. 2 and Figs. S3 and S4 of the Supplementary Material reveal the uniform distribution of the Pt particles with mean particle size of 3.5 ± 0.9 nm, 3.7 ± 1 nm and 3.2 ± 0.8 nm for samples with $x= 0.2, 0.3$ and 0.4 nominal Mo content, respectively. It can be concluded that the applied reduction-precipitation method leads to fine dispersion of the Pt particles. Within the experimental error, the Mo content has practically no effect on the average particle size and on the particle size distribution, either.

In good accordance with XRD result, electron diffraction patterns revealed the presence of Mo-doped rutile phase in the samples. The micrographs reveal the coexistence of a few large, faceted nanorod-like mixed oxide rutile crystallites (indicated by white arrows) with the carbon-containing composite material; their presence is attributed to the relatively high mixed oxide content of the samples, as in micrographs of a composite with 50 wt% mixed oxide – 50 wt% active carbon no such crystallites were observed [22, 23].

Diffraction patterns corresponding to the platinum crystallites also appear, although there is considerable overlap between features arising from rutile and fcc Pt (see Fig. 2C and Figs. S3C and S4C of the Supplementary Material). Additionally, it can be seen that platinum nanoparticles are deposited onto the surface of both the large mixed oxide crystallites and the composite material.

Energy-dispersive X-ray spectroscopy (EDX) measurement

The elemental composition of the platinum loaded electrocatalyst samples was evaluated by analyzing different regions by EDX technique; the average composition values are collected in Table 3.

Table 3. Composition of Pt loaded electrocatalyst samples measured by XRD, EDX and XPS in comparison with nominal values.

Samples ^{a)}	Mo _{subst.} , %	Ti/Mo (at/at)		(Ti+Mo+O)/C		Pt (wt%) EDX
		EDX	XPS	(wt/wt) EDX	(wt/wt) XPS	
Pt/Ti _{0.8} Mo _{0.2} O ₂ -C	20	83.7/16.3±1.7	77.3/22.7	68.9/31.1±1.8	55.5/44.5	20.1±0.6
Pt/Ti _{0.7} Mo _{0.3} O ₂ -C	25	78.4/21.6±3.8	73.7/26.3	66.1/33.9±1.3	58.6/41.4	21.9±0.4
Pt/Ti _{0.6} Mo _{0.4} O ₂ -C	28	74.7/25.3±0.8	62.6/37.4	64.5/35.5±2.3	57.4/42.6	19.2±0.8

^{a)} Expected composition of Ti_(1-x)Mo_xO₂ mixed oxide reflects desired Ti/Mo atomic ratio

Standard deviations in compositional values proved to be small, indicating that the atomic composition of the samples and the individual particles was highly homogenous. As far the Pt content is related, the nominal and measured values are in good agreement, confirming the success of the uniform deposition of the Pt nanoparticles on the Ti_(1-x)Mo_xO₂-C composite materials, as already demonstrated on the TEM images.

However, there is some increase in the Ti/Mo ratios when measured values are compared to nominal ones. This phenomenon is quite unexpected as Mo addition to the samples was a simple impregnation which was followed by a HTT. These steps would not lead to Mo release from the catalysts. Such a Mo loss can only be explained by acidic experimental conditions upon Pt loading. The decrease of the relative molybdenum content is presumably the consequence of partial dissolution of the less stable Mo species, not incorporated into the lattice, upon Pt deposition. Probably this Mo loss explains why the Ti_(1-x)Mo_xO₂/C weight ratios measured by EDX are also smaller than expected.

X-ray photoelectron spectroscopy (XPS) examination

In order to assess the surface composition of the investigated Pt/Ti_(1-x)Mo_xO₂-C electrocatalysts, XPS measurements were carried out on samples in the as prepared state (i.e. after 8 h annealing in Ar at 600°C). The observed composition data and the chemical states of the constituents are summarized in the following Table 4.

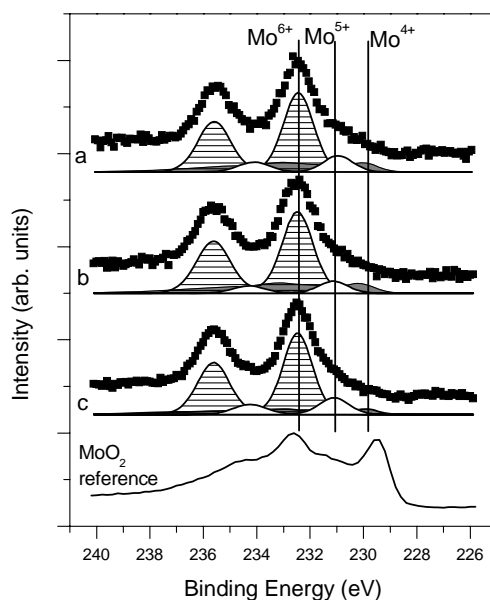


Fig. 3 Mo 3d XPS spectra of the Pt/Ti_{0.8}Mo_{0.2}O₂-C (a), the Pt/Ti_{0.7}Mo_{0.3}O₂-C (b), and the Pt/Ti_{0.6}Mo_{0.4}O₂-C (c) electrocatalysts, along with the reference spectrum of MoO₂. ■: Mo⁴⁺, □: Mo⁵⁺, ▨: Mo⁶⁺.

The ratio of the Ti- and Mo-content corresponds acceptably to the nominal composition in all three cases, which is in contrast to the findings of EDX. Actually, as XPS is more surface sensitive than EDX, the relative abundance of Mo as deduced from the XPS measurement suggests accumulation of non-incorporated Mo on the surface, probably in the form of less stable species. On the other hand, the oxygen content is always somewhat more than the nominal value (even with the assumption of Mo being completely in the 6+ oxidation state), suggesting the presence of excess surface oxygen species. Indeed, although the most intense contribution to the O 1s line occurs at a binding energy characteristic for TiO₂-related materials [31], the line shape is asymmetric towards higher binding energies, which can be related to the occurrence of surface hydroxyl groups and adsorbed water (spectra not shown). The carbon content of the samples is predominantly in the graphitic form; both EDX and XPS analysis indicate more carbon than the nominal value. In contrast to the EDX results, the concentration of Pt derived from XPS data is substantially more than the nominal 20 wt.%, which may be due to the well dispersed nature of Pt. Pt is predominantly metallic with a Pt 4f_{7/2} binding energy slightly above 71 eV. A weak contribution around 72.5 eV arises from Pt²⁺ species, which probably formed as a result of air exposure after synthesis.

Table 4. Summary of the XPS analysis of the 20Pt/Ti_(1-x)Mo_xO₂-C samples (the samples were used in the *as prepared* state).

Sample	Element (photoelectron peak, binding energy (eV), chemical state, concentration (atomic %))														
	Pt 4f _{7/2}			Mo 3d _{5/2}			Ti 2p _{3/2}			O 1s			C 1s		
	BE	Ch. state	Conc.	BE	Ch. state	Conc.	BE	Ch. state	Conc.	BE	Ch. state	Conc.	BE	Ch. state	Conc.
Pt/Ti _{0.8} Mo _{0.2} O ₂ -C	71.1	metal	6.5	229.9	Mo ⁴⁺	1.9	458.8	Ti ⁴⁺	6.5	530.2	M-ox	25.3	284.4	Gra- phite	59.8
	72.6	Pt ²⁺		231.1	Mo ⁵⁺		232.4	Mo ⁶⁺							
Pt/Ti _{0.7} Mo _{0.3} O ₂ -C	71.2	metal	6.3	230.0	Mo ⁴⁺	2.4	458.8	Ti ⁴⁺	6.7	530.3	M-ox	26.9	284.4	Gra- phite	57.7
	72.5	Pt ²⁺		231.1	Mo ⁵⁺		232.5	Mo ⁶⁺							
Pt/Ti _{0.6} Mo _{0.4} O ₂ -C	71.1	metal	5.8	229.9	Mo ⁴⁺	3.2	458.8	Ti ⁴⁺	5.5	530.2	M-ox	24.4	284.4	Gra- phite	61.1
	72.5	Pt ²⁺		231.1	Mo ⁵⁺		232.5	Mo ⁶⁺							

M-ox stands for metal oxide

According to the Ti 2p spectra, titanium is fully oxidized and is in the +4 oxidation state in all samples.

In Fig. 3 the Mo 3d spectra of the three samples are compared. All spectra are dominated by a strong peak pair located around 232.5 and 235.6 eV binding energy; these contributions arise from the Mo 3d_{5/2-3/2} spin-orbit doublet of Mo⁶⁺ (MoO₃) [32]. Apart from this, a pronounced asymmetry on the low binding energy side of the spectra indicate the presence of Mo ions in lower oxidation state as well.

According to the literature, the Mo⁵⁺ ionic state is expected to give a peak pair around 231 eV [32]. The Mo⁴⁺ ionic state, on the other hand, is characterized by a more complex 3d spectrum consisting of a combination of a narrower and a broad spin-orbit splitted doublet with the most intense contribution between 229-230 eV [33,34]; the two peak pairs represent differently screened final states. In this work the Mo⁴⁺ contribution was modeled by a line shape measured on a vacuum annealed MoO₂ powder sample (lowermost spectrum of Fig. 3). The Mo 3d spectra of the three samples were satisfactorily fitted with the Mo⁶⁺, Mo⁵⁺ and Mo⁴⁺ contributions, indicating that no metallic Mo species are on the surface. Modeling of the spectra suggests that for the Pt/Ti_{0.8}Mo_{0.2}O₂-C and the Pt/Ti_{0.7}Mo_{0.3}O₂-C samples roughly 20 % of the Mo content is in the 4+ state, while for the Pt/Ti_{0.6}Mo_{0.4}O₂-C catalyst this value decreases to around 10%.

From our previous work with Ti_(1-x)W_xO₂ mixed oxide based electrocatalysts we know that air exposure results in excessive oxidation of tungsten into the +6 oxidation state, regardless to the extent of the incorporation of the doping metal. Similarly, in the case of the Ti_(1-x)Mo_xO₂-based system we believe that air exposure induced oxidation of molybdenum leads to the dominance of the Mo⁶⁺ ionic state. The Mo⁴⁺ signal should mainly arise from Mo ions deeply incorporated into the rutile lattice, where they are protected from oxidation. The somewhat higher relative amount of strongly oxidized Mo species in the case of the Pt/Ti_{0.6}Mo_{0.4}O₂-C sample is thus in agreement with the higher amount of non-incorporated Mo deduced from the XRD data.

Electrochemical measurement

At first cyclic voltammograms measured on the Pt/Ti_(1-x)Mo_xO₂-C catalysts at the beginning (Fig. 4A) and the end (Fig. 4B) of the "pre-leaching" process are presented along with those obtained on the 20 wt% Pt/C reference.

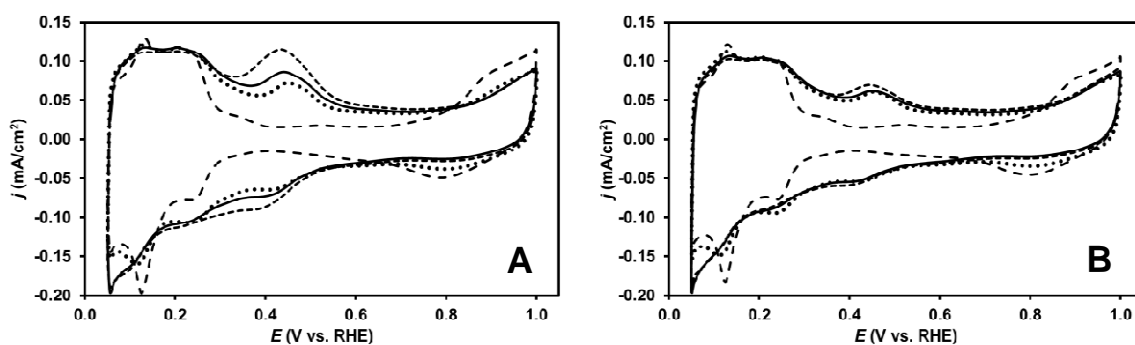


Fig. 4 Cyclic voltammograms of Pt/Ti_(1-x)Mo_xO₂-C catalysts obtained after the 1st (A) and the 2nd CO_{ads} stripping measurements (B). Dotted line: Pt/Ti_{0.8}Mo_{0.2}O₂-C; solid line: Pt/Ti_{0.7}Mo_{0.3}O₂-C; short dashed line: Pt/Ti_{0.6}Mo_{0.4}O₂-C. Results obtained on the parent Pt/C catalyst (dashed line) are given for comparison. Recorded in 0.5 M H₂SO₄ at 10 mV s⁻¹, T = 25 °C. The "pre-leaching" test for 30 cycles was carried out after the 1st and before the 2nd CO_{ads} stripping measurement at 100 mV s⁻¹.

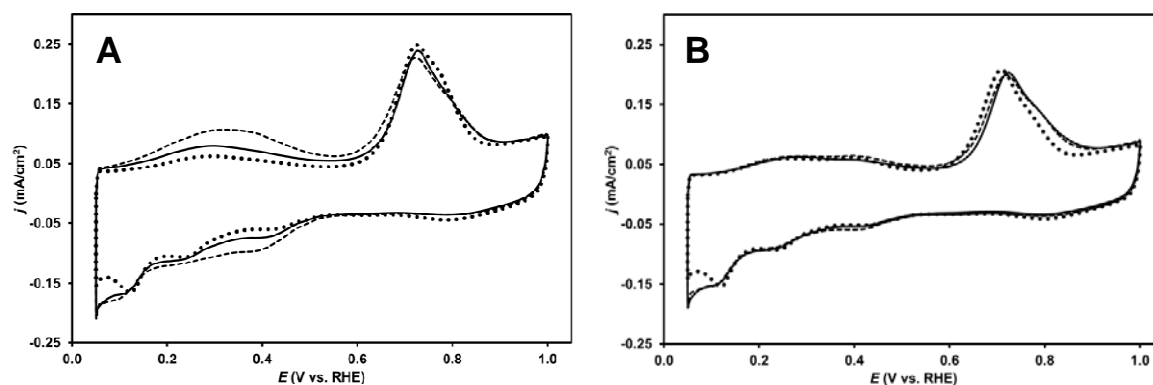
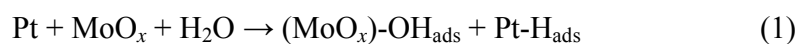


Fig. 5 Argon purged CO_{ads} stripping voltammograms of Pt/Ti_(1-x)Mo_xO₂-C catalysts before (A) and after (B) the “pre-leaching” procedure. Dotted line: Pt/Ti_{0.8}Mo_{0.2}O₂-C; solid line: Pt/Ti_{0.7}Mo_{0.3}O₂-C; dashed line: Pt/Ti_{0.6}Mo_{0.4}O₂-C. Recorded in 0.5 M H₂SO₄ at 10 mV s⁻¹, T = 25 °C. The “pre-leaching” was carried out for 30 cycles at 100 mV s⁻¹.

A typical voltammogram of Pt electrocatalysts with the classical features of the under-potentially deposited hydrogen adsorption/desorption between 50 mV and 350 mV was observed on all samples (Fig. 4.). Besides, redox peaks between 380 mV and 530 mV were additionally detected on Ti_(1-x)Mo_xO₂-C composite supported Pt catalysts. Similar results have been reported for Pt/Ti_{0.7}W_{0.3}O₂ electrocatalysts by Wang *et al.* [8]. An increased Coulombic charge of the hydrogen region was assigned to hydrogen spillover from Pt to the Ti_{0.7}W_{0.3}O₂ mixed oxide. As an analogy from the related art [35,36] the redox peak pair observed in the Pt/Ti_(1-x)Mo_xO₂-C samples can be interpreted as the intercalation/de-intercalation of H atoms into the MoO_x lattice resulting in the formation/decomposition of hydrogen molybdenum bronze like species.

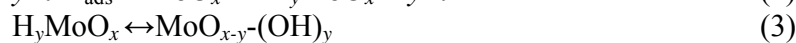
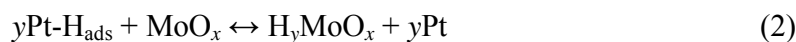
As shown in Fig. 4A the intensity of this redox peak pair depends on the Mo content in the samples and, as can be seen in Fig 4B, the decrease of the peak intensity was observed upon “pre-leaching”. The data suggest that the free, unincorporated MoO_x phase seems to have low resistance to synergic corrosion effect of acidic environment and cyclic polarization within a broad polarization window and may be dissolved upon aging.

According to the bifunctional mechanism the adsorbed hydroxyl species (OH_{ads}) on oxophilic molybdenum is responsible for increased CO tolerance of the Mo-containing catalysts at very low electrode potentials [7,37]. Formation of surface hydroxyls are shown schematically by equation 1:



This is considered as the reactive moiety with adsorbed CO [38].

Furthermore, in acidic solutions molybdenum hydrogen bronze H_yMoO_x could be formed by the “spillover” of hydrogen from Pt sites to Mo, that may have a contribution to enhanced catalytic activity toward CO electrooxidation [7,18]:



For such a process the close vicinity of the noble metal and MoO_x species is prerequisite. Results of TEM measurement proved atomic closeness of the $\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2$ mixed-oxide and Pt.

In our case one may expect that during the CV measurements intercalation of hydrogen into non-incorporated, segregated MoO_x oxide phase can result in a non-stoichiometric molybdenum hydrogen bronze H_yMoO_x formation [7]. Hou *et al.* [17] proposed that hydrogen and CO spillover occur on H_xMoO_3 , which leads to the facile removal of CO by the active water bound on the H_xMoO_3 in accordance with the bifunctional mechanism. Moreover the importance of hydrogen spillover from Pt to $\text{Ti}_{0.7}\text{W}_{0.3}\text{O}_2$ mixed oxides in the high activity for CO oxidation was demonstrated in Ref. [8]. Accordingly, in the CO_{ads} stripping measurements very good initial CO tolerance is expected, which should change as the unincorporated MoO_x species are dissolved during the “pre-leaching” treatment.

To investigate the influence of the mixed-oxide content of the support on the Pt catalyst in the presence of carbon-monoxide, Ar-purged CO_{ads} stripping voltammetry was performed using the different Pt/ $\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2$ -C samples (Fig. 5). The changes of the behavior of the Pt/ $\text{Ti}_{0.8}\text{Mo}_{0.2}\text{O}_2$ -C catalyst in Ar- and H_2 purged CO_{ads} stripping as a result of the “pre-leaching” process are compared to those observed for 20Pt/C and PtRu/C reference catalysts (Quintech) as well (Fig. 6).

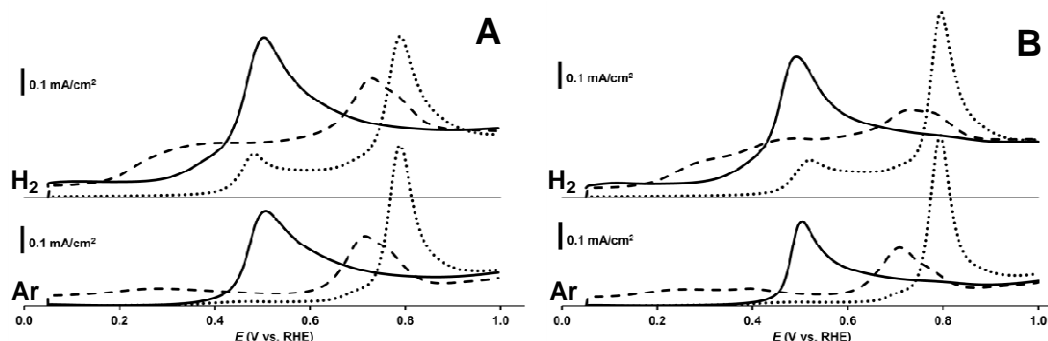


Fig. 6 The 1st (A) and the 2nd CO_{ads} stripping voltammograms (B) obtained after Ar purging and H_2 purging in the anodic scan on the Pt/ $\text{Ti}_{0.8}\text{Mo}_{0.2}\text{O}_2$ -C (dashed line), the PtRu/C (solid line) and the Pt/C Quintech (dotted line) catalysts. Recorded in $0.5 \text{ mol dm}^{-3} \text{ H}_2\text{SO}_4$ solution between 50 and 1000 mV potential limits with 10 mV s^{-1} sweep rate. Before the 2nd CO_{ads} stripping measurements the “pre-leaching” procedure was performed in Ar purged $0.5 \text{ mol dm}^{-3} \text{ H}_2\text{SO}_4$ solution between 50 and 1000 mV potential limits with 100 mV s^{-1} sweep rate.

Ar-purged CO_{ads} stripping results (Fig. 6) show that the onset potential corresponding to the CO oxidation reaction for Pt/ $\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2$ -C catalysts shifts toward less positive potential values in comparison to that obtained over the reference Pt/C sample. On the 20Pt/C catalysts the main CO_{ads} stripping peak is located at ca. 795 mV (see Table 5 and Fig. 6), while the corresponding peak in the case of the Pt/ $\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2$ -C catalysts is around 705-725 mV. Beyond the main peak two overlapping “pre-peaks” centered between ca. 200 mV and 400 mV can be seen over the Mo doped samples. These “pre-peaks” are typical for Pt-based catalysts containing Ti, Ta, W, and Mo suboxides, and are generally attributed to oxidation of weakly bonded CO [39,40]. The results indicate that the composite supported Pt catalysts, by means of the Mo species in close contact to Pt nanoparticles, could supply a source of hydroxyl intermediates

which are consumed during oxidation of CO to CO₂, cleaning this way the Pt active site from CO poisoning species [7]:

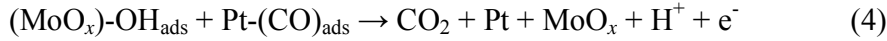


Table 5. Influence of the Ti/Mo ratio in Ti_(1-x)Mo_xO₂ mixed oxides on the electrochemical performance of the corresponding Pt/Ti_(1-x)Mo_xO₂-C electrocatalysts. Results obtained on the parent Pt/C and the PtRu/C catalysts are given for comparison.

Samples	E _{CO,onset} ^{a)} mV	E _{CO,max} ^{b)} mV	1 st ECSA _{CO} ^{c)} m ² /g _{Pt}	2 nd ECSA _{CO} ^{c,d)} m ² /g _{Pt}	Difference ^{e)} %
Pt/Ti _{0.8} Mo _{0.2} O ₂ -C	50 ^{f)}	705	72.9	58.1	20.3
Pt/Ti _{0.7} Mo _{0.3} O ₂ -C	50 ^{f)}	725	70.5	54.7	22.4
Pt/Ti _{0.6} Mo _{0.4} O ₂ -C	50 ^{f)}	715	73.4	55.8	24.0
PtRu/C	255	505	118.5	88.9	25.0
Pt/C	625	795	113.2	101.1	10.7

a) The onset potential for the CO electrooxidation;

b) The position of the main CO_{ad} stripping peak determined from the 2nd CO-stripping measurements;

c) 1st ECSA_{CO} and 2nd ECSA_{CO} values were calculated from the 1st and 2nd CO-stripping measurements, respectively; upon calculation the area of the “pre-peak” was not taken into account.

d) Before the 2nd CO-stripping measurements the “pre-leaching” test for 30 cycles at 100 mV s⁻¹ was performed;

e) $\{1 - (2^{\text{nd}} \text{ECSA}_{\text{CO}} / 1^{\text{st}} \text{ECSA}_{\text{CO}})\} \times 100\%$;

f) The E_{CO,onset} of the “pre-peak” observed on Pt/Ti_(1-x)Mo_xO₂-C catalysts.

On CO_{ads} stripping voltammograms after the “pre-leaching” process we observed the large decrease of the “pre-peak” mostly for samples Pt/Ti_{0.7}Mo_{0.3}O₂-C and Pt/Ti_{0.6}Mo_{0.4}O₂-C (Figure 5B). It appears that the current density values in this potential range became reduced to the level of the sample Pt/Ti_{0.8}Mo_{0.2}O₂-C. These observations support continuous removal/dissolution of Mo species from the surface and in the same time stabilization of the surface Ti/Mo ratio. One can suppose that free Mo species, i.e. those not incorporated into the rutile lattice, may be more sensitive to redox polarization than incorporated ones and thus may be removed when submitted to the “pre-leaching” procedure. These observations are in a good agreement with results presented in Ref. [41] obtained on the Pt_{0.8}Mo_{0.2} alloy and MoO_x@Pt core-shell anode electrocatalysts. It has been demonstrated [41] that even if the CO tolerance of the MoO_x@Pt system with MoO_x core and relatively low precious metal content is superior to Pt_{0.8}Mo_{0.2} alloy catalyst, this catalyst has shown poor durability due to leaching of the Mo. The conclusion that unstable unincorporated Mo species are responsible for the good initial CO tolerance is further supported by Ref. [42], where the “pre-peak” of the Pt/Ti₃O₅-Mo catalyst centered at ca. 450 mV was attributed to the faradaic process of intercalation/de-intercalation of H atoms into the MoO_y lattice. The disappearance of the Mo redox peak after 1000 cycles of cyclic polarization observed on this catalyst and decrease of the “pre-peak” intensity accompanied with simultaneous shift to higher potential in the CO stripping voltammograms were also demonstrated [42].

Accordingly, it can be stated that there are surface Mo species which survive the “pre-leaching” process and are quite stable. Considering the dependence of the extent of the changes of the CO_{ads} stripping behavior on the initial Mo content, these species can be identified as those incorporated into the rutile lattice. Apparently, when doped into TiO₂ at low enough level, the oxophilic doping metal (W, and Mo) is protected from dissolution [43]. At the moment, the nature of the oxygenated species provided upon H⁺ insertion (eq. (2)) and/or by chemisorption of

water molecules (eq (1)) over the Mo-doped rutile matrix is still unclear. Nevertheless, our data show that these protected Mo species may have crucial importance on CO electrooxidation and therefore on long term CO tolerance of Pt based electrocatalysts.

We tested the Pt/Ti_{0.8}Mo_{0.2}O₂-C catalyst with high degree of Mo incorporation by argon- and hydrogen-purged CO_{ads} stripping voltammetry before and after the “pre-leaching” process in comparison with reference commercial Pt/C and PtRu/C catalysts (see Fig. 6). Comparison of the CO_{ads} stripping behavior of the Ti_{0.8}Mo_{0.2}O₂-C electrocatalyst to the 20Pt/C catalyst (see Fig. 6) demonstrates increased CO tolerance of the composite supported catalyst. CO_{ads} stripping voltammograms measured on the Pt/Ti_{0.8}Mo_{0.2}O₂-C sample before and after the “pre-leaching” process (Fig. 6A and B) confirm the good stability of the catalyst in this potential range. According to the Ar-purged voltammograms, the onset potential for CO oxidation on the composite supported catalyst is not higher than 50 mV.

H₂-purged CO_{ads} stripping voltammetry reveals the potential at which vacancies are formed in the CO_{ads} monolayer [44, 45]. Once CO desorbs from a fraction of the catalytic sites, Pt sites become available for the dissociation/electrooxidation of H₂ molecules resulting in a steep increase of the current up to the diffusion-limited plateau [44-46].

The upper traces of Fig. 6A and 6B show the anodic current curves measured by H₂-purged CO_{ads} stripping voltammetry before and after the “pre-leaching” process. On the Pt/Ti_{0.8}Mo_{0.2}O₂-C catalyst, an increase of the current at around 200 mV is found, which means that the onset potential of hydrogen oxidation is shifted toward less positive potentials in comparison with Pt/C and PtRu/C (around 400 and 300 mV, respectively).

Ar- and H₂-purged CO_{ads} stripping experiments were also performed with electrocatalysts with higher nominal Mo content. The results are summarized in Table 5. The data indicate the very good CO tolerance of the Ti_(1-x)Mo_xO₂-C composite supported electrocatalysts. The substantial decrease of the electrochemically active surface area during the “pre-leaching” treatment indicates the transition from a mechanism for CO tolerance dominated by interaction between Pt and unincorporated Mo-oxide to one in which Mo ions stabilized by incorporation into the rutile lattice are relevant. The notable decrease of the ECSA_{CO} for PtRu/C during the “pre-leaching” may point to the sensitivity for dissolution even in this system.

Based on all the presented results, we can conclude that our catalysts supported on the Ti_(1-x)Mo_xO₂-C composite material perform better than commercial Pt/C or even PtRu/C. Our observations suggest that the optimal composition for the Pt/Ti_(1-x)Mo_xO₂-C electrocatalysts prepared by our method is around x=0.2-0.25.

Conclusion

As a continuation of our previous work, our aim was to investigate mixed titanium-molybdenum-oxide-carbon composite supported Pt electrocatalysts. Ti_(1-x)Mo_xO₂-C composite materials with three different Ti to Mo ratio were synthesized by a multistep sol-gel-based synthesis method followed by high-temperature heat treatment. According to the results of XRD measurements at the end of synthesis rutile phase was obtained with high degree of Mo incorporation. Nevertheless, incorporation in the case of samples containing more added molybdenum was found to be incomplete. It seems that incorporation of molybdenum under the used conditions has a maximum.

The 20 wt.% Pt loading was deposited onto the composite carrier via modified NaBH₄-assisted EG reduction-precipitation method. The results of EDX studies showed good agreement with the nominal composition of the 20 wt.% Pt loaded catalysts. A comparison of EDX and XPS data suggested surface enrichment of Mo and some Mo dissolution during Pt loading.

Enhanced initial CO tolerance of the Pt electrocatalysts prepared by using the composite support materials were evidenced by the appearance of a CO-oxidation related “pre-peak” and by considerable shift of the maximum of the main CO oxidation peak towards less positive potential compared to the reference commercial Pt/C. The onset potential of H₂ oxidation also shifted to lower potential in comparison with the reference catalyst. According to the results of electrochemical measurements there is a good correlation between the Mo content of the samples and the extent of CO tolerance.

However, the difference in the electrochemical behavior of the samples with different Ti/Mo ratio gradually disappeared as a result of the “pre-leaching” process. The amount of unincorporated Mo being present in the sample at the beginning dissolved during the “pre-leaching” treatment, so the extent and nature of the Pt-Mo interaction changed. From these results it can be concluded that although the high initial Mo concentration results in significant Pt-Mo interaction the unincorporated Mo is not stable and in long term use only the incorporated content is available and relevant.

Acknowledgements

This work was supported by the National Development Agency [grant number KTIA_AIK_12-1-2012-0014]. Financial support by the OTKA-project [grant numbers K100793 (Zoltán Pászti) and K112034 (István Bakos)] is greatly acknowledged.

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Supplementary material for

Effect of Mo incorporation on electrocatalytic performance of Ti-Mo mixed oxide-carbon composite supported Pt electrocatalysts

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1. Details of the preparation of the $\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2\text{-C}$ support materials

The main features of the optimized preparation procedure of $\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2\text{-C}$ composite support materials are briefly summarized in Fig. S1.

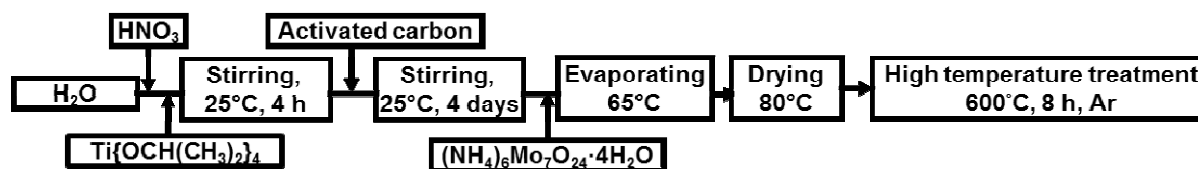


Fig. S1 Flow chart for preparing $\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2\text{-C}$ composite materials by using multistep synthesis route.

As the first step the calculated amount of titanium-isopropoxide was added dropwise to aqueous solution of HNO_3 , into a round bottom flask with continuous agitation to obtain transparent TiO_2 sol. The system was then stirred for 4 hours in order to initiate formation of rutile phase nuclei [S1] and then 250 mg of carbon black (CABOT, Black Pearls 2000) suspended in 5 ml water was added to the mixture. The rutile phase nanoclusters are supposed to precipitate onto the surface of the carbon and the crystal growth takes place within the suitable aging period. In this study the synthesis mixture was stirred for 4 days in order to form TiO_2 -rutile phase on the surface of the carbon. After the 4 days aging the given amount of the ammonium-molybdate-tetrahydrate was added to the synthesis mixture and was heated up to 65°C in order to evaporate the solvent under continuous stirring. Finally, the powder was dried in an oven overnight at 80°C . The obtained solid has been submitted to heat treatment in order to provoke isovalent cation substitution of Ti^{4+} in the TiO_2 -rutile lattice by Mo^{4+} , i.e. incorporation of molybdenum. The heat treatment of the samples was carried out in a vertical quartz tube in a tubular furnace. The samples were heated up to 600°C in argon flow and were kept at this temperature for 8 hours, then the system was cooled down to room temperature also in argon flow.

2. Details of the preparation of the 20 wt.% Pt/ $\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2\text{-C}$ electrocatalysts

$\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2\text{-C}$ composite materials were loaded with 20 wt.% Pt via a modified NaBH_4 -assisted ethylene-glycol (EG) reduction-precipitation method (see Fig. S2) [S1-S3].

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In all cases 0.643 mmol $\text{H}_2\text{PtCl}_6 \cdot 6 \text{H}_2\text{O}$ was solved in 50 ml of ethanol and 200 mg of the samples was suspended in the solution. A solution prepared by mixing of 7.8 mmol NaBH_4 and 3.7 ml EG was added dropwise to the suspension at 65 °C with continuous stirring. After 3 hours of stirring at 65 °C, 15 ml 0.2 M HCl was added to the suspension and stirred for an additional 2.5 hours at room temperature in order to deposit the Pt particles onto the support material. The materials were washed three times with 50 ml water and filtered by centrifugation in order to remove the chloride ions and dried at 80 °C in an oven overnight.

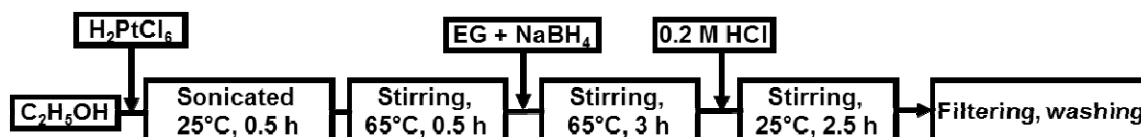


Fig. S2 Flow chart for synthesis of $\text{Pt}/\text{Ti}_{(1-x)}\text{Mo}_x\text{O}_2\text{-C}$ electrocatalysts.

3. Transmission electron microscopy (TEM) investigations

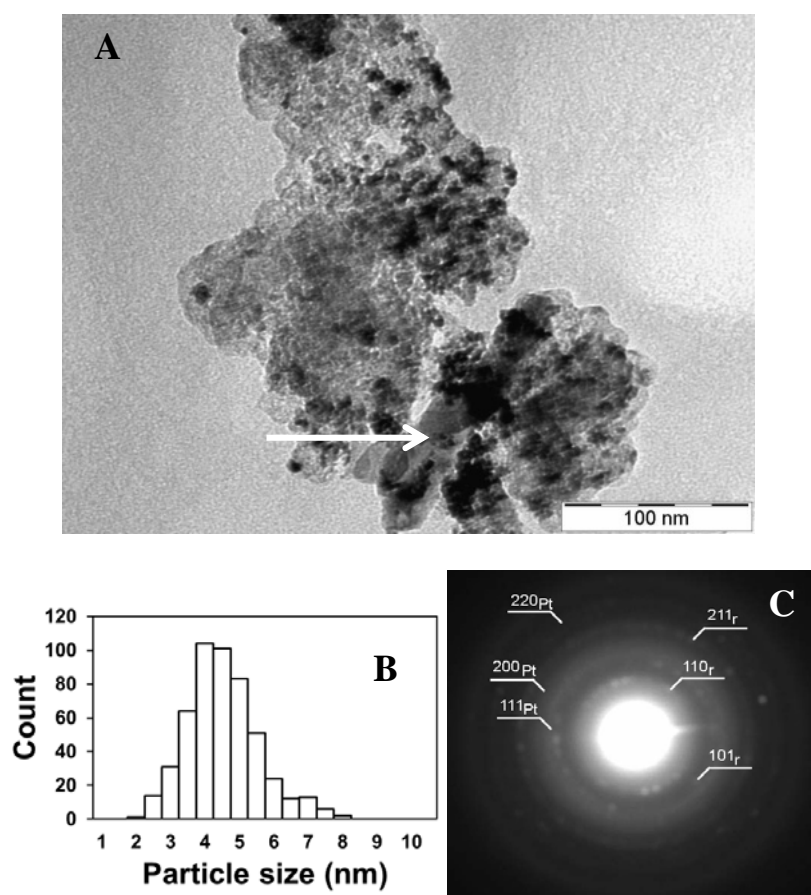


Fig. S3 TEM image (A), histogram of the Pt particle size distribution (B) and electron diffraction pattern (C) of the $\text{Pt}/\text{Ti}_{0.7}\text{Mo}_{0.3}\text{O}_2\text{-C}$ electrocatalyst; “r” stands for the Mo-doped rutile phase.

TEM micrograph, electron diffraction pattern and particle size distribution of the 20 wt.% $\text{Pt}/\text{Ti}_{0.8}\text{Mo}_{0.2}\text{O}_2\text{-C}$ catalyst are shown in Figure 2 of the main text. Very similar morphologies were observed for the $\text{Pt}/\text{Ti}_{0.7}\text{Mo}_{0.3}\text{O}_2\text{-C}$ and the $\text{Pt}/\text{Ti}_{0.6}\text{Mo}_{0.4}\text{O}_2\text{-C}$ catalysts; micrographs,

electron diffraction patterns and histograms for particle size distribution for these samples are presented in Figs. S3 and S4.

Fig. 2 of the main text and Figs. S3 and S4 reveal the uniform distribution of the Pt particles with mean particle size of 3.5 ± 0.9 nm, 3.7 ± 1 nm and 3.2 ± 0.8 nm for samples with $x = 0.2$, 0.3 and 0.4 nominal Mo content, respectively. It can be concluded that, within the experimental error, the Mo content has practically no effect on the average particle size and on the particle size distribution, either. The Pt particles are relatively small. The applied reduction-precipitation method leads to fine dispersion and accordingly high active surface area is expected in electrochemical tests. In good accordance with XRD result, electron diffraction patterns revealed the presence of the Mo-containing rutile phase in the samples. Patterns corresponding to the platinum crystallites also appear, although there is considerable overlap between features arising from rutile and fcc Pt (see Fig. 2C of the main text and Figs. S3C and S4C).

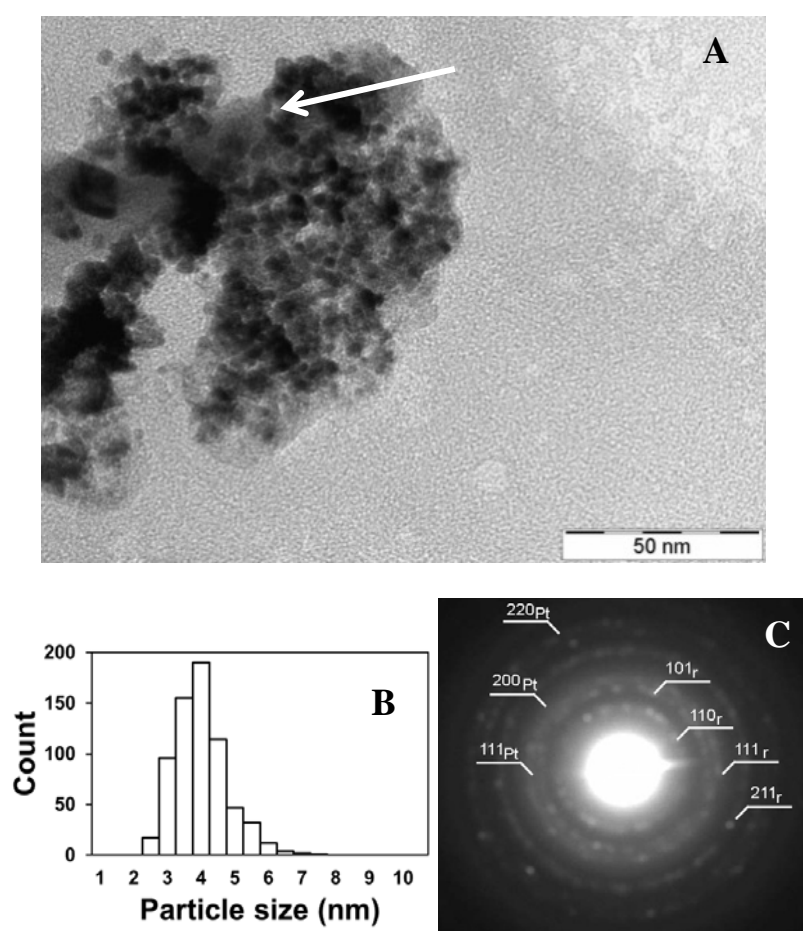


Fig. S4 TEM image (A), histogram of the Pt particle size distribution (B) and electron diffraction pattern (C) of the Pt/ $\text{Ti}_{0.6}\text{Mo}_{0.4}\text{O}_2\text{-C}$ electrocatalyst; “*r*” stands for the Mo-doped rutile phase.

Representative images for average spots of the samples reveal the coexistence of a few large, faceted nanorod-like Mo-doped rutile crystallites (indicated by white arrows) with the carbon-containing composite material. In the micrographs of a composite with 50 wt% mixed oxide – 50 wt% active carbon no such crystallites were observed [S1]; their presence is attributed to the increased mixed oxide content of the samples. It can be seen that platinum

nanoparticles are deposited onto the surface of both the large mixed oxide crystallites and the composite material.

4. Electrochemical characterization before and after the “pre-leaching” procedure

It is known that less stable transition metal oxides tend to dissolve under acidic conditions and high potentials. Such leaching and subsequent uncontrolled re-deposition of certain components of the electrocatalyst may be detrimental for the performance of the fuel cell, so stability against initial leaching should be checked. Accordingly, in our work data from the very first CV scans are analyzed and summarized in Fig. S5 for the Pt/Ti_{0.8}Mo_{0.2}O₂-C, the Pt/Ti_{0.6}Mo_{0.4}O₂-C and the commercial 20 wt.% Pt/C Quintech (C) catalysts.

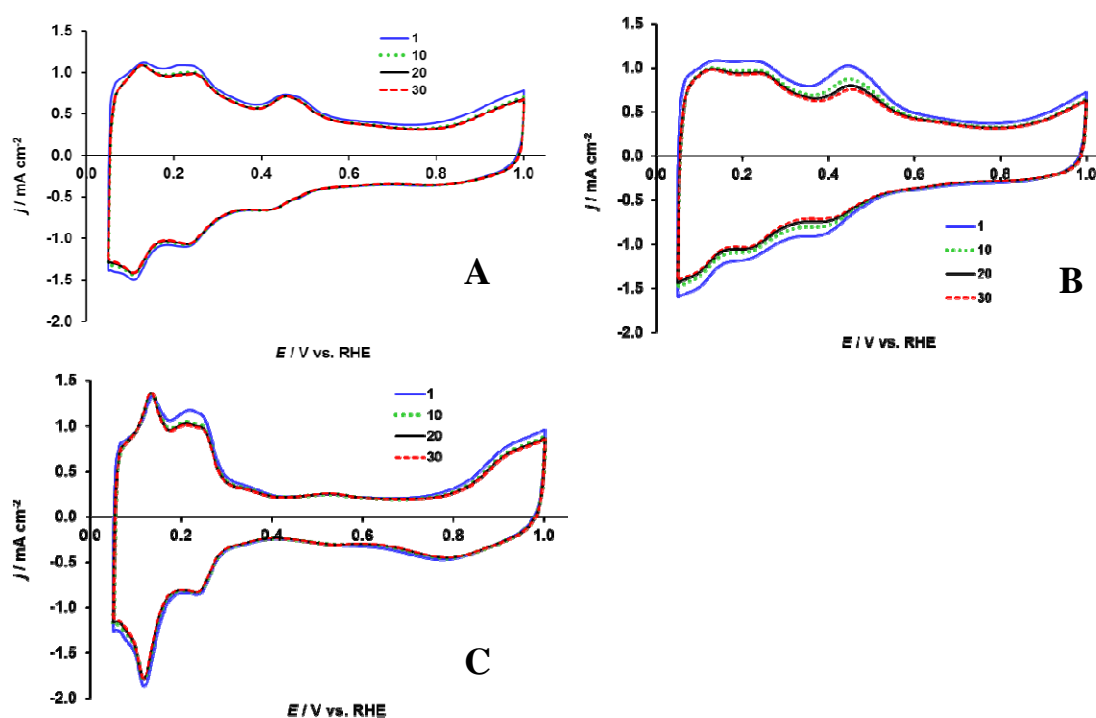


Fig. S5. Cyclic voltammograms of the Pt/Ti_{0.8}Mo_{0.2}O₂-C (A), the Pt/Ti_{0.6}Mo_{0.4}O₂-C (B) and the commercial 20 wt.% Pt/C Quintech (C) catalysts. CVs recorded in the 1st cycle (blue curve), 10th cycle (dotted green line), 20th (black curve) and 30th cycle (dashed red line) of the “pre-leaching” test in Ar-saturated solution after the 1st CO_{ads} stripping measurements. Recorded in Ar purged 0.5 mol dm⁻³ H₂SO₄ solution between 50 and 1000 mV potential limits with 100 mV s⁻¹ sweep rate.

In the case of Pt/Ti_{0.8}Mo_{0.2}O₂-C relatively small changes can be seen during the first 30 CV scans. The substantial changes of the voltammograms of Pt/Ti_{0.6}Mo_{0.4}O₂-C can be explained by dissolution or rearrangement of the non-incorporated Mo species during the first cycles. The results of our work demonstrate that these initial changes are more expressed on the samples containing the higher amount of non-incorporated Mo species, thus we can expect better leaching properties from catalysts with less non-incorporated Mo-oxide (Fig. S5.A).

In addition, it may be noted that the air exposed catalyst surface may be covered by carbonaceous contaminants or may be heavily oxidized, so its properties are clearly different from the one formed after some cyclic potential scans. In fact, in order to reach a quasi-stable steady state of the surface, a cleaning or conditioning stage consisting of a certain number of CV scans is routinely applied prior to collection of electrochemical data. Indeed, even in the

case of the commercial Pt/C sample a few CV cycles are needed for stabilization of the surface conditions.

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