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Direct Observation of Layered-to-Spinel Phase Transformation in Li_2MnO_3 and Spinel Structure Stabilised after the Activation Process†

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Abstract

Li_2MnO_3 is an important parent component in the lithium- and manganese-rich layered oxides (LMRs), which are one of the promising positive electrode materials for the next-generation lithium ion rechargeable batteries. Here, we report the layered-to-spinel phase transformation in Li_2MnO_3 during the initial charging process to characterise its unique delithiation behaviour, which gives an insight into the relationship between the structure, superior capacities and degradation of the LMR electrodes. The atomic-scale observation using scanning transmission electron microscopy (STEM) techniques suggests that the structural transformation occurs in a

biphasic manner within a single particle. The formed phase has a Li-defect spinel structure, indicating that the delithiation leads to Mn migration from the transition-metal layer to the Li layer, accompanied with some oxygen release. This layered-to-spinel phase transformation is an essential bulk process in the initial activation of Li_2MnO_3 . During the lithiation in the 1st discharging, the Mn remigration occurs and the layered structure is again formed with significant disordering. During the multiple cycles, the defect spinel structure is stabilised and becomes more oxygen-deficient with lower Mn valence. As a consequence, the amount of inserted Li decreases, which corresponds to the capacity and voltage fading observed in Li_2MnO_3 and the LMRs.

† Electronic supplementary information (ESI) available. See ***.

Keywords: Lithium-ion rechargeable batteries, Li_2MnO_3 , scanning transmission electron microscopy, electron energy loss spectroscopy.

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1. Introduction

Lithium ion rechargeable batteries (LIBs) have been widely used as a power source for portable devices. On the other hand, its wide applications to electric vehicles (EVs) further demand high power, high energy density, long life, low cost, and safe batteries. The most successful positive electrode materials are layered transition-metal oxides, LiCoO_2 and its derivatives, with their electrochemical capacities of ca. 150 mA h g^{-1} up to a practical potential of 4.2 V vs. Li/Li^+ .¹ Lithium- and manganese-rich layered oxides (LMRs), commonly described as a composite $x\text{Li}_2\text{MnO}_3 \cdot (1-x)\text{LiMO}_2$ ($M = \text{Ni, Co, etc.}$), have been extensively studied as a preferred candidate material for the next-generation LIB positive electrodes because they show high reversible capacities beyond 200 mA h g^{-1} .²⁻⁹ It is well known that LMRs show a characteristic irreversible voltage plateau at ca. 4.5 V vs. Li/Li^+ in the initial charging process. This plateau is considered to come from the activation of the nominally inactive Li_2MnO_3 component in the materials.^{3,5} Therefore, understanding the delithiation mechanism of the end-member component Li_2MnO_3 should be an important clue to elucidate the origin of the superior capacities of LMRs and also the origin of the deterioration after the multiple charge–discharge cycles. This information will be helpful to improve their electrochemical performance.

Li_2MnO_3 , also described as $\text{Li}[\text{Li}_{1/3}\text{Mn}_{2/3}]\text{O}_2$, has a layered structure with the space group of $C2/m$.^{10,11} Li ions occupy not only the Li layer but also the transition-metal (TM) layer with intralayer ordering between Li and Mn ions, giving the theoretical capacity of 459 mA h g^{-1} . The electrochemical profile of Li_2MnO_3 has a voltage plateau at ca. 4.5 V vs. Li/Li^+ in the initial charge and shows a large irreversible capacity in the subsequent discharge.¹² There have been some controversial reports about the charge compensation mechanism on the delithiation from Li_2MnO_3 . Firstly, the increase in Mn valence from +4 to +5 had been speculated for charge compensation,¹³ but the subsequent reports suggested that the oxidation state of Mn ions remains

unchanged based on the Mn K-edge X-ray absorption fine structure (XAFS) spectra.^{14–16} Alternatively, Dahn et al. proposed a simultaneous lithium and oxygen removal from the structure.¹⁷ The oxygen removal seems consistent with a large irreversible capacity observed in the initial charge–discharge cycle. Yu et al. showed a significant amount of O₂ release from the electrode, but they also pointed out that it merely accounts for about a half of the observed capacity in Li₂MnO₃.¹⁴ On the other hand, the first-principles calculations have indicated that the oxidation of oxygen anion (O²⁻/O⁻ redox couple) without O₂ release can explain the characteristic voltage plateau of Li₂MnO₃.^{18,19} Then, the formation of the peroxo-like species were experimentally suggested in the chemically-delithiated Li₂MnO₃ after the important studies provided by Tarascon and coworkers, who used the O 1s X-ray photoelectron spectroscopy (XPS) to examine the oxidation states of oxygen anions in the charged and discharged Li-rich layered oxide electrodes.^{20–22} Also, we have suggested the redox contribution of oxygen anions in the Li₂MnO₃ bulk structure and its reversible behaviour on charge–discharge cycles by using the O K-edge XAFS spectroscopy.²³

On the other hand, structural phase transformation during the delithiation process has been still unclear because the reported X-ray diffraction (XRD) profiles showed significant decreases in diffraction intensity, indicating a severe lattice disordering by the lithium extraction,¹⁴ which makes structural characterisation difficult. Some studies reported weak diffraction peaks attributable to spinel structure in the partially delithiated Li₂MnO₃.^{24,25} Recent advances in scanning transmission electron microscopy (STEM) techniques have offered a chance to gain the atomic column images for the delithiated/relithiated electrode materials.^{26–28} Wang et al. observed the Mn migration from the TM layer to the Li layer in Li₂MnO₃ charged to 4.8 V in the initial cycle, which was considered as a direct evidence of the formation of spinel-like domain during delithiation.²⁹ They further reported that the Mn occupancy in the Li layer disappeared in the

relithiated sample. Also, the spinel domains were observed at the particle surface in the 10th cycled discharged Li₂MnO₃ electrode sample, whereas the *C2/m* lattice structure was preserved in the bulk region.^{30,31} These studies concluded that the formation of spinel phase was a consequence of structural degradation and highly localised at the particle surface and the limited region in the bulk. Croy et al. investigated the structural evolutions of Li and Mn local environments in between the pristine Li₂MnO₃ and electrode samples charged to 5.0 V and discharged to 2.0 V by using solid-state nuclear magnetic resonance (NMR) and extended X-ray absorption fine structure (EXAFS) analyses.¹⁶ They reported that the lithium extraction from both the Li and TM layers and the subsequent reinsertion into both the layers for the charged and discharged samples, respectively. A substantial reduction of Mn valence from Mn⁴⁺ to Mn³⁺ with the Jahn-Teller distorted coordination environments was also suggested in the discharged material. Although these studies have provided important suggestions about the structural changes of Li₂MnO₃ on the initial delithiation/relithiation process, some details are still open to debate. The important mysteries are the process how the Li ions are deintercalated from the lattice (single-phase or two-phase reaction) and the crystal structure after the full extraction of Li ions in Li₂MnO₃. In this article, we report the phase transformation occurring in the Li₂MnO₃ lattice structure during the initial charge–discharge cycle and propose the two-phase delithiation mechanism on the charging process on the basis of the high angle annular dark field (HAADF) STEM observation coupled with electron energy loss spectroscopy (EELS). The lattice structures of the active materials in the 20th discharged and 21st charged electrodes are also characterised, which relates the degraded structure to the capacity and voltage fading during the prolonged electrochemical operation.

2. Experimental

Li_2MnO_3 was synthesised by a solid-state reaction at 700 °C for 12 h with the starting materials $\text{LiOH}\cdot\text{H}_2\text{O}$ (Wako Pure Chemical Industries) and MnCO_3 (Kojundo Chemical Laboratory) in a molar ratio of 2.1 : 1.0. The details are described elsewhere.³² The obtained particle size was estimated to be ca. 2 μm on average, which consisted of small crystallites of ca. 100 nm in diameter. The chemical composition was confirmed from the inductively coupled plasma-atomic emission spectrometry (ICP-AES, ICPS-8100, Shimadzu), and the average Mn oxidation state was estimated from the iodometric titration measurements.

A positive electrode was prepared from a mixture of Li_2MnO_3 , acetylene black (Denki Kagaku Kogyo), and polyvinylidene difluoride (PVDF, Kureha) in a weight ratio of 80 : 10 : 10, which was spread onto an aluminium foil with 1-methyl-2-pyrrolidone (NMP) and then dried at 80 °C under vacuum overnight. The electrode was pressed to a typical thickness of 35–40 μm . A foil of metallic lithium (0.2 mm in thickness, > 99.9%, Honjo Metal) was used as counter and reference electrodes. These components were assembled together with the Celgard 2500 separator and soaked in the electrolyte solution, which were sealed in an aluminium-coated laminate-type cell in an Ar-filled glove box. The electrolyte solution was 1M LiPF_6 solution dissolved in anhydrous ethylene carbonate (EC) and ethylmethyl carbonate (EMC) with a volumetric ratio of 3 : 7 (Kishida Chemical).

Electrochemical measurements were performed at 50 °C on an automatic cycling and data recording system (HJ1001SD8, Hokuto Denko). The cells were galvanostatically cycled between 4.8 and 2.0 V vs. Li/Li^+ at a current rate of 23 mA/g. The cells were carefully disassembled at desired states of charge/discharge in the glove box and rinsed with dimethyl carbonate (DMC) to remove the electrolyte solution residue.

Synchrotron radiation powder XRD (SR-XRD) measurements were carried out for the

disassembled samples sealed in 0.5 mm ϕ borosilicate glass capillaries in the 2θ range of $4.2^\circ - 77.8^\circ$ with $\Delta\theta$ of 0.003° at the BL28XU beamline at SPring-8 (Hyogo, Japan). The incident radiation of $\lambda = 0.500 \text{ \AA}$ and two-dimensional X-ray detector PILATUS (DECTRIS) was applied. The acquired 2θ datasets were converted to ones corresponding to the Mo $K\alpha$ wavelength ($\lambda = 0.709 \text{ \AA}$).

^6Li MAS NMR spectra were acquired on a DD2 600 spectrometer at a magnetic field of 14.1 T with a wide-bore T3 MAS probe (Agilent Technologies). The powdered samples were packed into 1.2 mm ϕ MAS ZrO_2 rotors with airtight caps, which were spun at a spinning rate of 50 kHz during the experiments. The practical temperatures of spinning samples at 50 kHz were estimated to be ca. 60°C due to frictional heating. A rotor-synchronised Hahn echo pulse sequence ($\pi/2-\tau-\pi-\tau-\text{acq.}$) was used with the $\pi/2$ pulse width of $0.9 \mu\text{s}$ and a relaxation delay of 0.1 s. All spectra were referenced to 1M LiCl solution at 0.0 ppm.

Atomic column images were obtained by using the Cs-corrected scanning transmission microscope JEM-ARM200F (JEOL) operated at 200 kV. The samples were thin-sliced to less than 100 nm in thickness with focused ion beam (FIB, nanoDUET NB5000, HITACHI High-Technologies) and placed on a copper mesh and then transferred into the microscope. All these treatments were performed under vacuum or Ar atmosphere. The high angle annular dark field (HAADF) images were collected with the probe current of 8 pA and the collection semiangle of $42-168 \text{ mrad}$. The spatial resolution of the microscope is defined as ca. 0.1 nm with the electron beam convergence semiangle of 14 mrad . Electron energy loss spectroscopy (EELS) and energy dispersive X-ray spectrometry (EDS) analyses were carried out on GIF Quantum-ER energy filter (Gatan) and JED-2300T (JEOL), respectively, which were equipped on the microscope. The EELS mapping measurements were performed with the probe current of 22 pA

and the collection semiangle of 71 mrad. The exposure time was 0.05 s per pixel (1 pixel = 2×2 nm² in dimension) for Li and 0.25 s for O and Mn, respectively. The energy resolution was 1.25 eV.

3. Results and discussion

3.1. Phase transformation in Li₂MnO₃ during the 1st charge–discharge cycle

Fig. 1 shows the charge-discharge profile of Li₂MnO₃ for the 1st cycle. The sampling points for structural analyses are marked on the figure and also listed in Table S1. An irreversible voltage plateau is observed at ca. 4.5 V on the charging process. The charging and discharging capacities delivered at 50 °C are 442 and 214 mA h g⁻¹, respectively. Almost the full extraction of Li ions is achieved by the electrochemical measurements at high temperature.

Fig. 2 shows the SR-XRD profiles for the 1st cycle. The pristine Li₂MnO₃ sample can be indexed with the space group of *C2/m* and the refined lattice parameters of $a = 4.9259$ Å, $b = 8.5283$ Å, $c = 5.0116$ Å, and $\beta = 109.125^\circ$ (Fig. S1a†).³² During the charging to 4.8 V, all the diffraction intensities are monotonously decreased and broadened, indicating that the delithiation from Li₂MnO₃ results in a highly disordered material. Close inspection on the strongest 001 diffraction peak reveals that a shoulder peak appears and grows at higher diffraction angle with increasing the charging depth to 230 and 345 mA h g⁻¹, and finally the diffraction peak settles at the highest angle of $2\theta = 8.83^\circ$ for the sample charged to 4.8 V (Fig. 2, inset). This suggests that the delithiation proceeds in a two-phase reaction mechanism. The previous studies reported the appearance of the diffraction peaks attributable to cubic/tetragonal spinel phases for the charged samples.^{24,25} Some diffraction peaks observed in the present study (Sp. #5) are close in position to those for Li-poor cubic spinel Li _{δ -0.05}Mn₂O₄ (λ -MnO₂, Fig. S1b†),³³ although they are also similar to those of the P3-type stacking (*R-3m*) structure.³⁴ The electrode sample charged to 4.8 V shows

almost no residual peaks of the pristine Li_2MnO_3 , being consistent with the observed charging capacity of 442 mA h g^{-1} (relative to the theoretical capacity, 459 mA h g^{-1}).

When Li ions are reinserted into the material with the voltage down to 2.0 V, the diffraction peaks are close to the original positions for Li_2MnO_3 but with the reduced intensities (Sp. #7). This suggests that the material partly comes back to Li_2MnO_3 structure with low crystallinity. It is noteworthy that about a half of Li ions should be in the material discharged to 2.0 V according to the observed discharging capacity (Table S1), but its diffraction intensities are lower than those for the sample charged to 230 mA h g^{-1} (Sp. #3).

Fig. 3 shows the ^6Li MAS NMR spectra for the 1st cycle. The pristine Li_2MnO_3 electrode sample shows three isotropic shifts (Sp. #1). The minor sharp signal at ca. 0 ppm comes from the impurity phase such as Li_2CO_3 . The predominant signal at 728 ppm with a shoulder at ca. 760 ppm is assigned to the Li ions in the Li layer ($4h$ and $2c$ in Wyckoff position, respectively) and the other at 1476 ppm is attributable to the TM layer ($2b$ in Wyckoff position) in Li_2MnO_3 .^{35–37} On the charging process, the peak positions remain unchanged, but the signal intensities decrease monotonously. The Li ions are deintercalated from both the Li and TM layers (Fig. 3, inset). The sample disassembled at 4.8 V shows almost no signal, being consistent with the observed charging capacity. These results can be reasonably explained by the two-phase delithiation model between Li_2MnO_3 and Li-poor phase(s). Although the Li-poor phase(s) cannot be identified from the ^6Li NMR technique, the Li-poor cubic spinel $\text{Li}_{\delta-0.05}\text{Mn}_2\text{O}_4$ is invoked based on the above XRD result and the HAADF-STEM results as discussed below. On the discharging process, some of Li ions are reinserted into both the Li and TM layers in Li_2MnO_3 . Therefore, some domains of the Li-poor spinel phase revert to the pristine Li_2MnO_3 structure. This is consistent with the XRD result. We here notice that in the ^6Li NMR spectrum of the sample discharged to 2.0 V (Sp #7) the signal intensities are significantly reduced compared to the Li content estimated from the

electrochemical measurements (Fig. 3, inset). The actual Li content in the sample discharged to 2.0 V was confirmed by ICP-AES to be almost identical to the one from the electrochemical measurement. All the reintercalated ${}^6\text{Li}$ signals should be observed (semi-) quantitatively in the spectrum, even if the material is highly disordered, because the ${}^6\text{Li}$ signals for the known Li-Mn-O materials (LiMn_2O_4 , $\text{Li}_4\text{Mn}_5\text{O}_{12}$, $\text{Li}_2\text{Mn}_2\text{O}_4$, and layered LiMnO_2) are in the chemical shift range between 0 and 2000 ppm.^{38–40} We also observed an increase of the signal at ca. 0 ppm for the sample discharged at 2.0 V, which is attributed to the solid electrolyte interphase (SEI) components such as Li_2CO_3 and LiF .^{41,42} However, it does not fully account for the difference in Li content between the NMR and electrochemical measurements (Fig. 3, inset), even if the spectrum is acquired with the long relaxation delay of 20 s. Therefore, we speculate that the apparent signal loss is due to severe broadening of ${}^6\text{Li}$ signal by the strong dipolar interaction between Li ion and paramagnetic spins of Mn ions in the discharged material.

Previous studies have reported that the Li^+/H^+ exchange occurs as a major charge compensation mechanism on the electrochemical measurements at 55 °C, which is associated with the decomposition of electrolyte solvent on the electrode.^{34,43,44} We observed a small ${}^1\text{H}$ signal at 299 ppm, which can be assigned to the intercalated proton in the material during the charge (Fig. S2†).⁴³ However, its intensity is very small compared to the ${}^1\text{H}$ signal from the PVDF, the present study suggests that the Li^+/H^+ exchange seems less significant even at 50 °C.

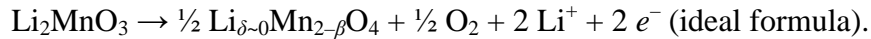
To confirm the two-phase delithiation model from the atomic scale resolution, we obtained the EELS and HAADF-STEM images for the electrode disassembled at the 50% SOC (230 mA h g^{-1}). Fig. 4 shows the EELS intensity mapping images for Li, O K- and Mn L-edges. The particle thickness was roughly estimated to be 50–100 nm from the ratio of the total electron beam intensity and zero-loss intensity. This particle was confirmed not to overlap with the other particles. The inhomogeneous Li distribution strongly supports the coexistence of Li-rich and

Li-poor regions within a single particle. The relatively homogeneous O and Mn intensity distribution indicates that the Li inhomogeneity does not come from the inhomogeneous particle thickness. The Li mapping also indicates that the Li extraction is not consistent with a simple core-shell or domino-cascade type reaction front models. From the electron energy loss (EEL) spectra it is found that the chemical states of Mn and O ions are different between the Li-rich and Li-poor regions (Fig. 4e–g). The Mn L- and M-edge spectra for the Li-rich region are similar in shape and peak top position to those for MnO₂, where Mn ions are tetravalent. The L-edge intensity ratio L_3/L_2 , which is used as a criterion of the oxidation state of TM ions,^{45–47} is also close to that for MnO₂. The fast Fourier-transformed (FFT) patterns of the HAADF image show diffraction spots indexed with $C2/m$, indicating that the Li-rich region remains as Li₂MnO₃ (Fig. 5b, inset and S3†). On the other hand, the Li-poor region shows the Mn L- and M-edge spectra similar to those for Mn₂O₃. The L_3/L_2 ratio is also close to that for Mn₂O₃. This suggests that the Li-poor region has a reduced Mn valence state close to trivalent. The FFT image indicates that the Li-poor region consists of cubic ($Fd-3m$) and/or tetragonal ($I4_1/amd$) spinel phase(s) (Fig. 5d, inset and S3†). These results are consistent with our preliminary EELS and selected area electron diffraction (SAED) measurements on the TEM images acquired for the Li₂MnO₃ electrode disassembled at 4.8 V (Fig. S4b,d†). Also, the O/Mn atomic ratio in the Li-poor region is found to be smaller (by ca. 2.4) than that in the Li-rich region, based on the O K-edge and Mn L-edge EELS intensity ratio (See also Fig. S5a†). Although the chemical composition determined from the EELS intensity is still less accurate, this strongly suggests that some oxygen atoms are removed from the material to be the Li-poor spinel phase during the delithiation process.

The HAADF-STEM image for the Li-rich region shows dumbbell-like atomic lines (-Li-Mn-Mn-) characteristic of TM layer in Li₂MnO₃ (Fig. 5c). It also shows the stacking faults along c axis as suggested by the XRD profile for the pristine material.³² There is no clear

evidence of the cation mixing within the TM layer and the Mn migration into the Li layer, although our previous report on the pristine material indicated the cation mixing within the TM layer.³² It is possible to speculate that the delithiation easily proceeds from the cation mixing regions, which then transform into the spinel-like framework. The HAADF image for the Li-poor region is smoky in appearance and it is difficult to observe the atomic column lines (Fig. 5d). This would be related to the significant lattice disordering and is consistent with the severe intensity decrease in the XRD profiles on delithiation (Fig. 2). However, on closer inspection the Li-poor region shows the Mn migration into the Li layer (Fig. 5e), which corresponds to the formation of spinel-like structure.^{29–31}

We summarise the phase transformation of Li_2MnO_3 on the initial charging process. We propose that the delithiation from Li_2MnO_3 proceeds in a two-phase reaction;



Recently, the two-phase delithiation mechanism was suggested for Li_2MnO_3 by using Raman spectroscopy and principle component analysis.⁴⁸ In that study, a MnO_2 -type phase was suggested in the charged material, whereas the other Raman spectroscopy study considered the formation of spinel-type structure.²⁵ Here, we emphasise that the resultant spinel phase is not a surface degradation product during the charging process. It is also noteworthy that the chemical composition, especially the O/Mn ratio, of the spinel phase is not well fixed, and it is most likely defective. From the EEL spectra, it is found that the Li-poor spinel phase involves the trivalent Mn ions. The Mn L- and O K-edge XAFS spectra also suggested the reduced Mn ions and some contribution of the oxygen in lattice to charge compensation (Fig. S6†). Although the Mn reduction on the charging process is considered to be a consequence of the charge compensation for oxygen vacancy formation,⁴⁹ it may also be associated with the partial oxidation of the lattice oxygen atoms. These results imply that the Li-poor spinel phase, tentatively expressed as

$\text{Li}_{\delta\sim 0}\text{Mn}_{2-\beta}\text{O}_4$, observed on the delithiation from Li_2MnO_3 is different from the Li-poor spinel $\text{Li}_{\delta\sim 0}\text{Mn}_2\text{O}_4$ observed on the delithiation from LiMn_2O_4 (Fig. S6†). Similar discussion was done by Phillips et al., who considered the Mn-substoichiometric $\text{Li}_x\text{Mn}_{4/3}\text{O}_4$ spinel in the discharged material.³¹ The recent first-principles calculations suggested that the spinel-like structure is stabilised compared to the pristine layered structure for Li_xMnO_3 with low Li contents.^{30,50,51} Therefore, the layered-to-spinel transformation of $\text{Li}_{\delta\sim 0}\text{MnO}_3$ will be thermodynamically plausible. This is the activation mechanism during the 1st charging process accompanying the Mn migration within the structure and the oxygen loss from the structure, where the latter may be kinetically hindered to some extent, resulting in the highly disordered defective structure different from $\text{Li}_{\delta\sim 0}\text{Mn}_2\text{O}_4$. The discharging process is further complicated. The XRD and NMR results suggest that the reaction expressed above appears partially reversible. The SAED patterns for the samples charged to 4.8 V and discharged to 2.0 V at room temperature indicated that the cubic spinel phase transforms back into the monoclinic layered phase (Fig. S4†). Therefore, it is expected that the Mn remigration from the Li layer to TM layer occurs and overall crystal structure reverts to the layered structure during the relithiation in the 1st cycle. This is consistent with the annular bright field (ABF) STEM observation by Wang et al.²⁹ On the other hand, the average oxidation state of the Mn ions is close to +3 to +3.5 in the discharged state (Fig. S4d and S6†). Considering the above information and the fact that the discharging capacity reaches only about a half of the charging capacity (Fig. 1), the discharged material consists of layered LiMnO_2 with a small fraction of Li_2MnO_3 , or may be described as partially-disordered $\text{Li}[\text{Li},\text{Mn}]\text{O}_2$.

3.2. Structural characterisation of Li_2MnO_3 after the 20th charge–discharge cycle

The crystal structures of the active material after the 20th charge–discharge cycle are also examined. Fig. 6 shows the charge–discharge profiles for the 1st, 2nd, 5th, 10th, 15th, and 20th

cycles. They show the capacity and voltage fading up to the 20th cycle. The profiles are almost identical after the 10th cycle, suggesting that the structural changes are almost completed within the first 10 cycles. The charging and discharging capacities at the 20th cycle are 133 and 130 mA h g⁻¹, respectively. The dQ/dV curves and the charge–discharge capacities as a function of cycle number are also shown in Fig. S7†.

The SR-XRD profiles of the cycled materials are shown in Fig. 7. The diffraction profiles for the 2nd charged and discharged samples (Sps. #10,11) are similar to the ones for the 1st charged and discharged samples (Sps. #5,7), but some peaks become much broader and are significantly reduced in intensity. After the 20th charge–discharge cycle, the discharged sample shows a severely broadened profile (Sp. #12), and the main peak positions are reasonably close to those of LiMn₃O₄ (Fig. S1†).⁵² The 21st charged sample also shows a broad profile similar to that of the 20th discharged one, but the peak positions are slightly shifted to higher diffraction angles (Sp. #13). Therefore, the charged sample is considered to be a Li-poor Li_{δ-0}Mn₃O₄-like phase. This suggests that the pristine layered structure is transformed and gradually stabilised into a spinel-like framework during the multiple delithiation/relithiation processes. This is also supported by the ⁶Li MAS NMR spectra, where the Li environment in the pristine Li₂MnO₃ structure is completely lost and a new one probably associated with a Mn³⁺-bearing phase is stabilised in the 20th cycle (Fig. S8†).¹⁶

To understand the structural origin of the electrochemical fading behaviour, the HAADF-STEM images are acquired for the electrode disassembled at 2.0 V in the 20th cycle and that at 4.8 V in the 21st cycle (Fig. 8). Both the discharged and charged materials show the FFT patterns representative of the cubic ($Fd-3m$) and/or tetragonal ($I4_1/amd$) spinel phases (Fig. 8b and e, inset). The atomic column images for both the materials are smoky in appearance, but in several regions the Mn migration into the Li layer representative of the spinel structure can be

identified. The Mn migration is observed in the bulk regions as well as at the surface regions. Therefore, we conclude that both the discharged and charged materials are completely transformed into the spinel-like structure in the 20th charge–discharge cycle. The previous STEM studies reported that the Li₂MnO₃ structure was preserved in the bulk but changed to the spinel and cubic rock-salt structures at the surface for the Li₂MnO₃ particle in the electrode disassembled at 2.0 V after 10 cycles at room temperature.^{30,31} Such a difference among the present and previous studies can be partly attributed to the temperature at which the electrochemical measurements were performed. The charge–discharge cycling at 50 °C would accelerate the irreversible layered-to-spinel phase transformation in the structure.

The EELS intensity mapping indicates the inhomogeneous Li distribution in the 20th discharged material (Fig. 9b), suggesting the heterogeneous relithiation in the degraded material. Surprisingly, the EEL and XAFS spectra indicate that both the discharged and charged materials hold the divalent and trivalent Mn ions (Fig. 9e,g,h,j and S9†). Also, they show the lower O/Mn ratios of ca. 1.8 and 1.1, respectively (Fig. S5b and c†).

The previous studies reported that the layered LiMnO₂ with the space group of *C2/m* shows the layered-to-spinel transformation during the cycles.^{53,54} Therefore, we conclude that the layered LiMnO₂-like structure, which is considered to occur at the end of the initial cycle, is again transformed into the Li-poor spinel structure in the subsequent charging process, and finally the Mn ions are frozen at the Li layer, the layered structure is lost, and the spinel-like framework is stabilised both in the charged and discharged states after multiple cycles. The proposed phase transformation process is summarised in Fig. 10. The spinel phases in the charged and discharged materials have highly disordered and defective structures, and their compositions may be close to Li_{δ-0}Mn₃O₄ and LiMn₃O₄, respectively. It should be noted that the reported charge–discharge profiles of layered LiMnO₂ and LiMn₃O₄ are different from the present

curves,^{53,55} the structure, composition, and charge compensation mechanism of the cycled material should be more complicated. These findings suggest that the capacity and voltage fading in Li_2MnO_3 during multiple charge–discharge cycles are attributable to the structural stabilisation into the defect spinel phase both in the charged and discharged states, where the lithium reinsertion is suppressed by the structural reorganisation accompanying the oxygen removal and the lower-valence Mn redox couple ($\text{Mn}^{2+}/\text{Mn}^{3+}$) is activated.

4. Conclusions

We investigated the delithiation process of Li_2MnO_3 during the initial charge–discharge cycle and showed a direct evidence of the layered-to-spinel phase transformation in a two-phase reaction manner on the initial charge. The EELS/HAADF-STEM techniques provide the comprehensive information on elemental distribution, valence state, crystal structure, and visualisation of atomic column images, which indicated the two structural domains consisting of Li_2MnO_3 and Li-poor defect spinel phase in a single particle in the electrode sample at the 50% SOC. Therefore, the electrochemical activation of Li_2MnO_3 involves the formation of the spinel phase, which is associated with the oxygen loss from the pristine structure and the spontaneous Mn migration into the vacated Li layer site with the partial oxidation of the lattice oxygen and the reduction of Mn valence. We emphasise that the layered-to-spinel transformation occurring on the initial charge is not a consequence of the structural degradation, but it is the essential activation process of Li_2MnO_3 . Also, it is confirmed from the ^1H NMR that the Li^+/H^+ exchange is negligible on the electrochemical operation at 50 °C. Although the defect spinel structure would revert back to the layered structure during the initial relithiation, the detailed structure is now different from the pristine Li_2MnO_3 . It was also found that after the 20th cycle the spinel structure is stabilised in the fully-discharged/charged materials. Such stabilisation of the spinel framework compared to the

layered structure on Li reintercalation is considered as the structural degradation where the $\text{Mn}^{2+}/\text{Mn}^{3+}$ redox couple is activated and a limited amount of Li ions are reintercalated due to the irreversible oxygen loss from the lattice. This corresponds to the capacity and voltage fading in the charge–discharge profile of Li_2MnO_3 . We believe that the same fading behaviour that has been reported in LMRs is also caused by the layered-to-spinel transformation and its stabilisation during the charge–discharge cycles in the Li_2MnO_3 -rich domain in LMRs.

Author Contributions Statement

The manuscript is prepared through contributions of all authors: K.S. and M.O. designed the basic part of this research. T.M. prepared the pristine sample. K.S, O.M. and T.M. performed the SR-XRD measurements. K.S. and M.M. performed the solid-state NMR measurements. K.S. and M.O. planned the STEM measurements. K.Y. performed the soft X-ray XAFS measurements. K.S., M.O., T.M., M.M., K.Y., H.A., and T.O. contributed to the scientific discussion of these results. Y.Uch., T.O., E.M., and Z.O. supervised the research. K.S., M.O., M.M., H.A., and Y.Uk. wrote the manuscript, and all authors contributed to the manuscript reviewing and revisions.

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Figures and Figure captions

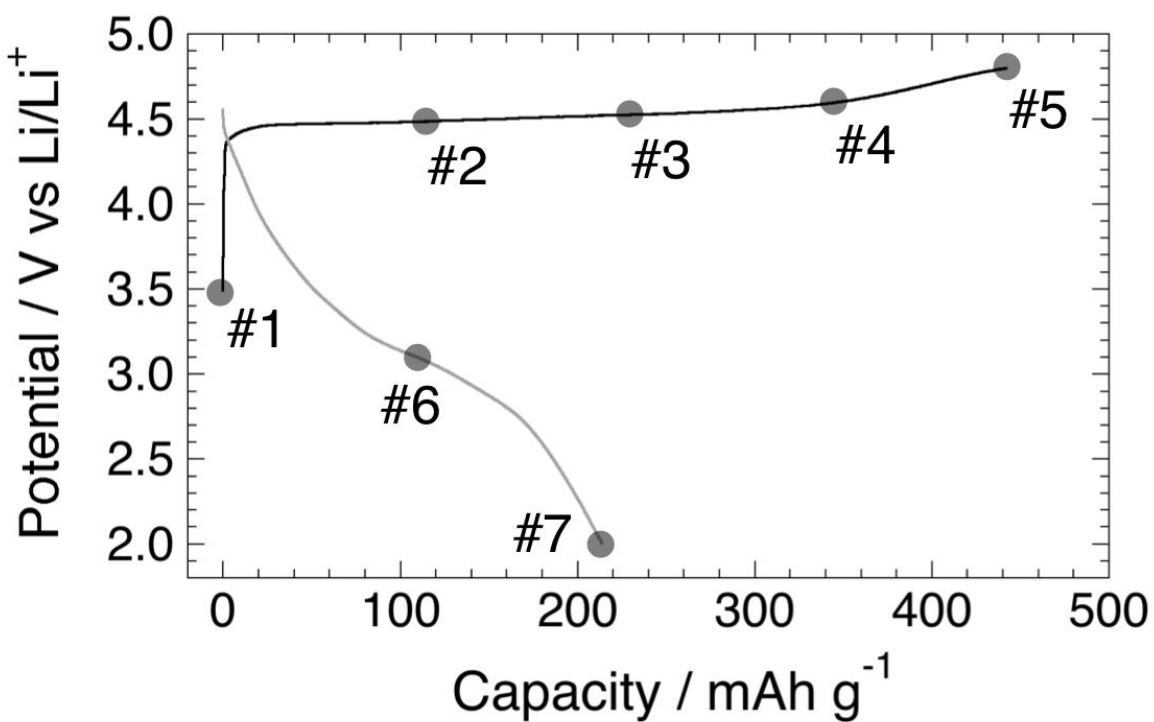


Fig. 1. Charge–discharge profiles of the Li/Li₂MnO₃ cell for the 1st cycle. The electrochemical measurements were performed at 50 °C between 4.8 and 2.0 V vs. Li/Li⁺ with a constant current of 23 mA/g. The sampling points are marked on the figure and listed in Table S1.

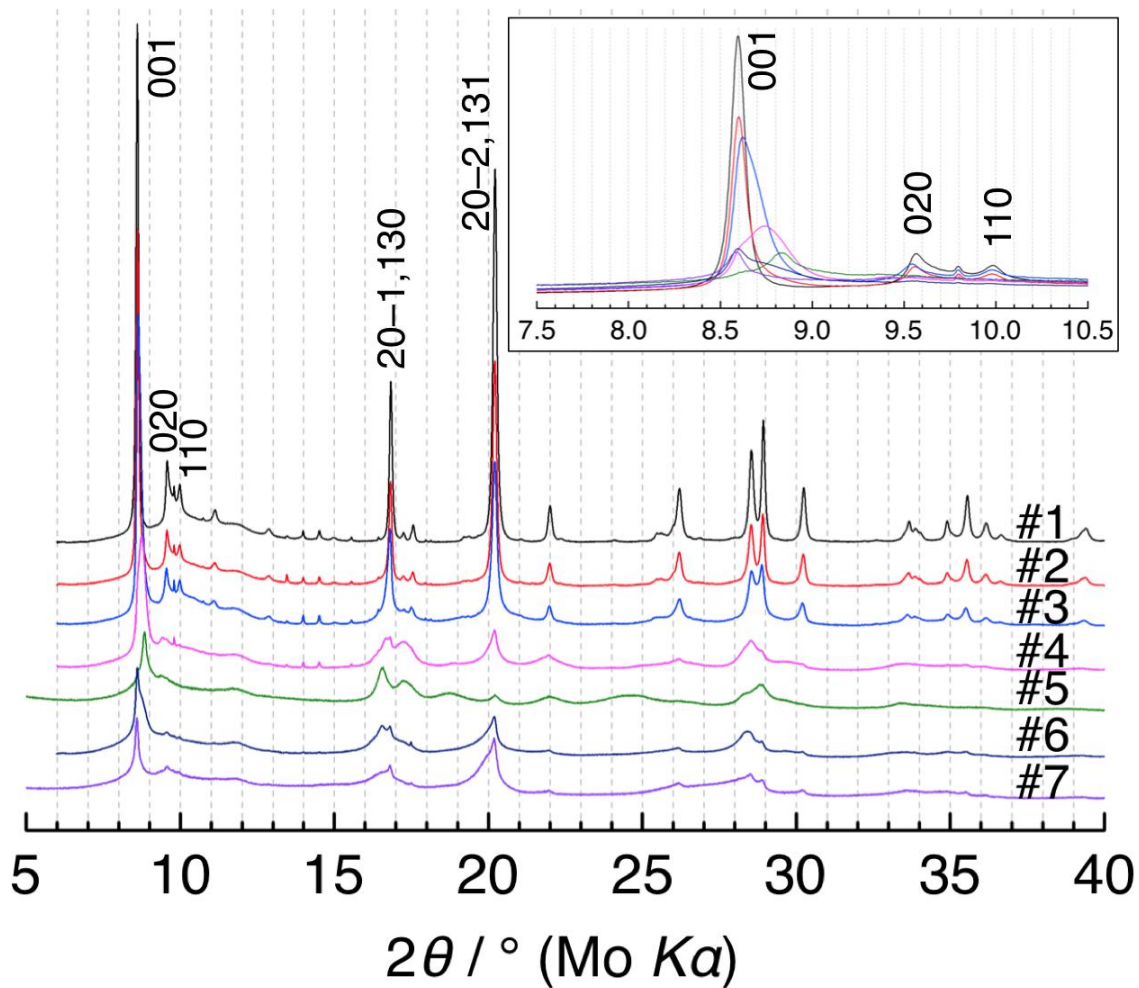


Fig. 2. SR-XRD profiles of the Li_2MnO_3 electrode for the 1st cycle. The wavelength is converted to that of Mo $K\alpha$ radiation. The strongest 001 diffraction peak is enlarged in the inset. The sample numbers are described in Fig. 1.

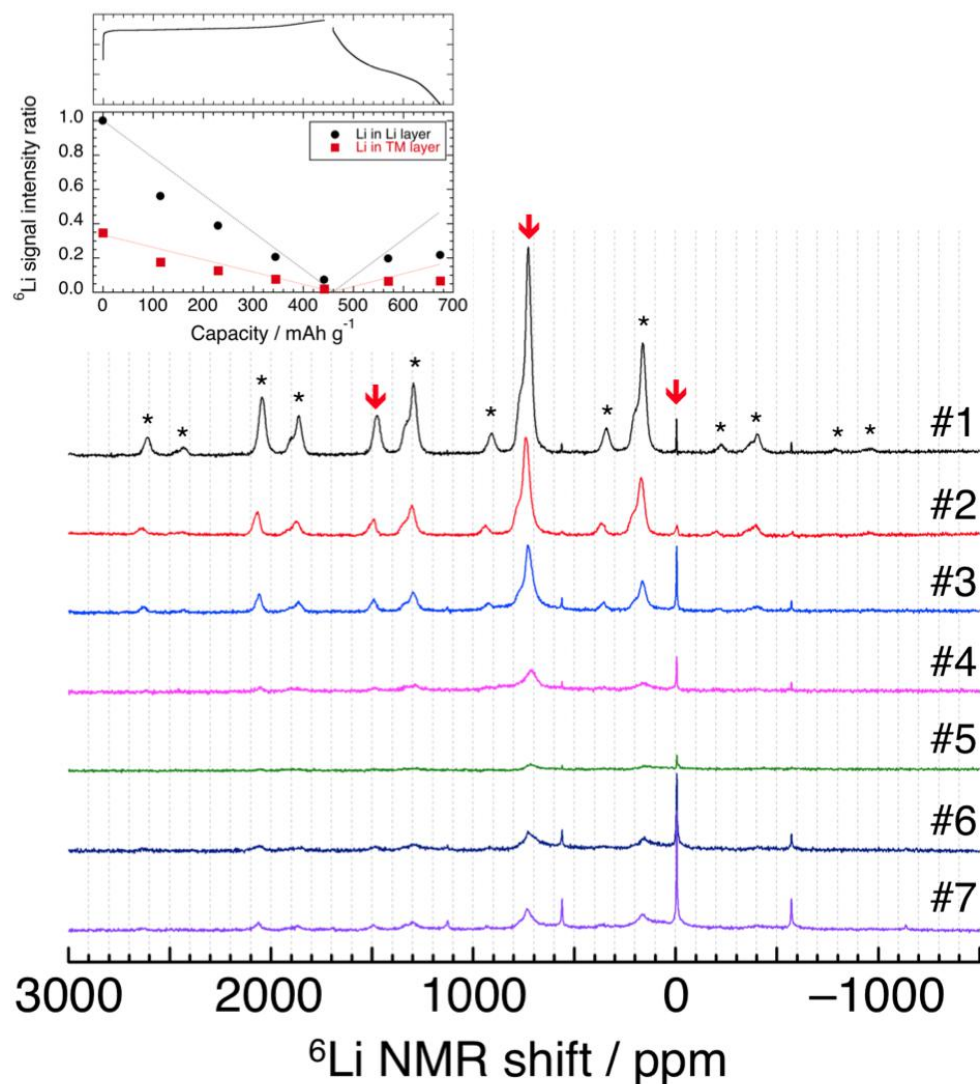


Fig. 3. ${}^6\text{Li}$ MAS NMR spectra of the Li_2MnO_3 electrode for the 1st cycle. The sample numbers are described in Fig. 1. The signal intensities are normalised to the scan numbers and sample weights in the rotors. Isotropic signals and their spinning sidebands are marked with arrows and asterisks on the spectrum #1, respectively. Intensity variation of the ${}^6\text{Li}$ NMR signals along with the potential profile is shown in the inset. The ${}^6\text{Li}$ signal in the Li layer is normalised to 1.0 in the pristine $\text{Li}[\text{Li}_{1/3}\text{Mn}_{2/3}]\text{O}_2$. The signal intensities are integrated including sideband manifolds. The straight lines are guides to the eye.

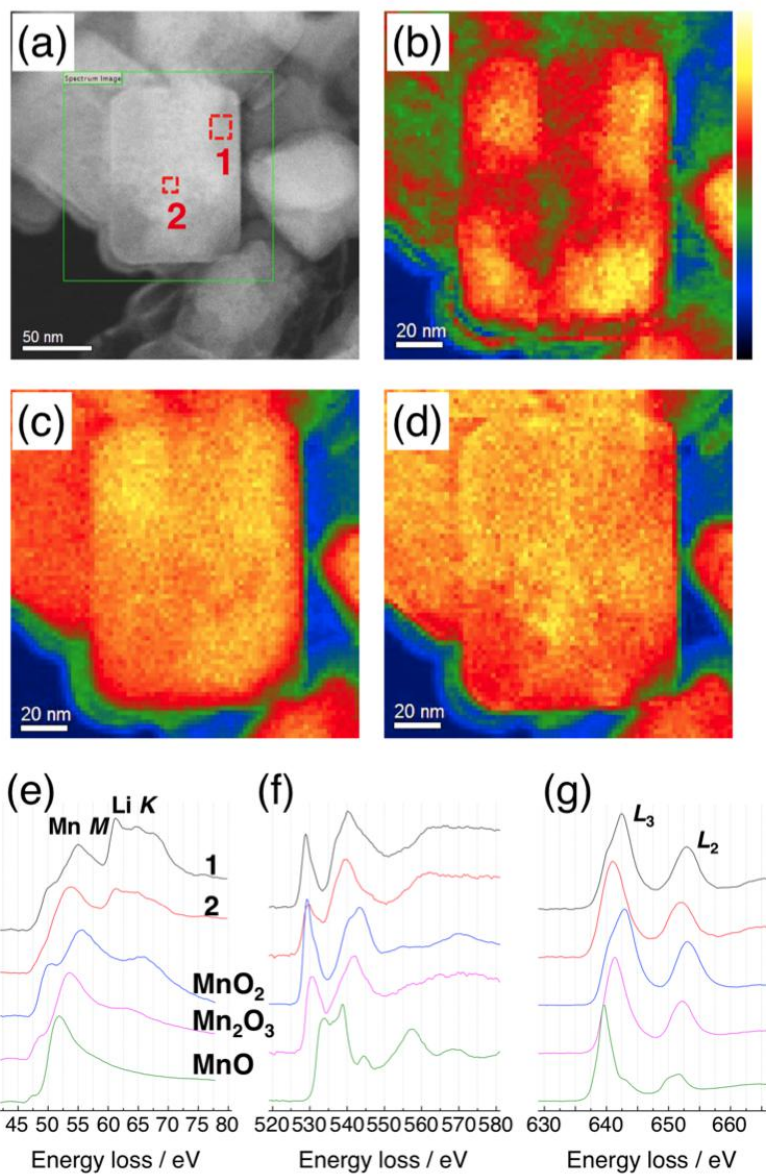


Fig. 4. EELS mapping images and EEL spectra for the electrode sample disassembled at the 50% SOC in the 1st cycle. (a) HAADF-STEM image, (b) Li K-edge, (c) O K-edge, and (d) Mn L-edge intensity maps. The EEL spectra at (e) Mn M-edge + Li K-edge, (f) O K-edge, and (g) Mn L-edge. The spectra 1 and 2 are acquired from the regions 1 and 2 (corresponding to the Li-rich and Li-poor regions) in (a), respectively. The reference spectra of MnO₂ (Mn⁴⁺), Mn₂O₃ (Mn³⁺), and MnO (Mn²⁺) are also shown.

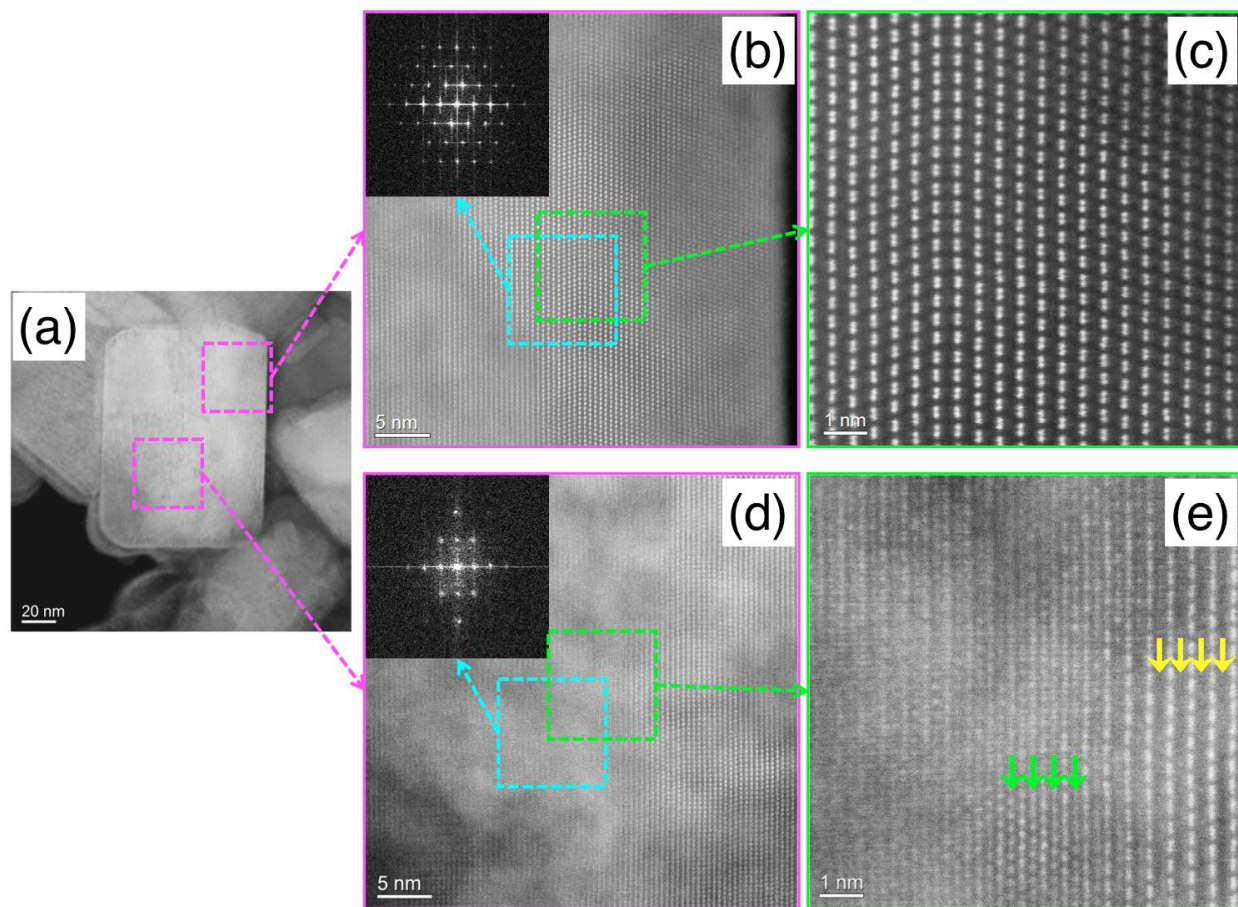


Fig. 5. HAADF-STEM images acquired along the $[110]_{\text{monoclinic}}$ zone axis for the electrode sample disassembled at the 50% SOC in the 1st cycle. (a) HAADF image of a selected single particle, (b,c) HAADF image of a selected area corresponding to the Li-rich region and (d,e) Li-poor region, respectively. The FFT patterns of the HAADF images are shown in (b) and (d). The Li and Mn atoms in the Li layers are marked with yellow and light green arrows, respectively (light Li atoms, if present, are invisible in the HAADF image).

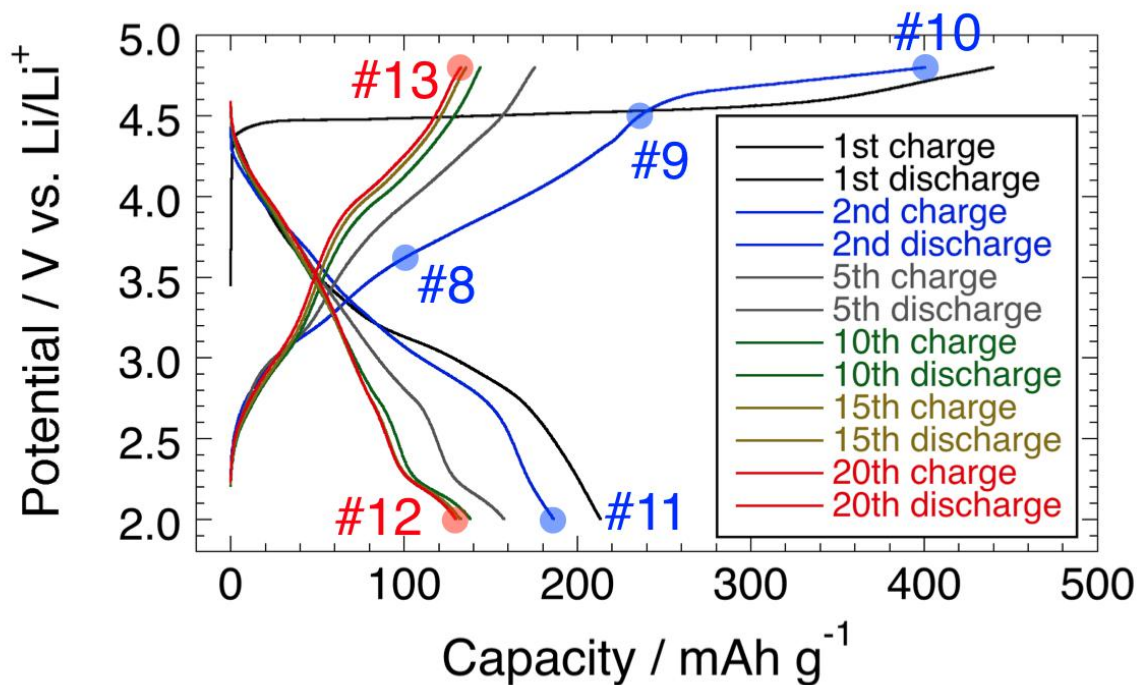


Fig. 6. Charge–discharge profiles of the Li/Li₂MnO₃ cell for the 1st, 2nd, 5th, 10th, 15th, and 20th cycles. The electrochemical measurements were performed at 50°C between 4.8 and 2.0 V with a constant current of 23 mA/g. The sampling points are marked on the figure and listed in Table S1 (The sample #13 for the electrode charged to 4.8 V in the 21st cycle is marked on the 20th charge).

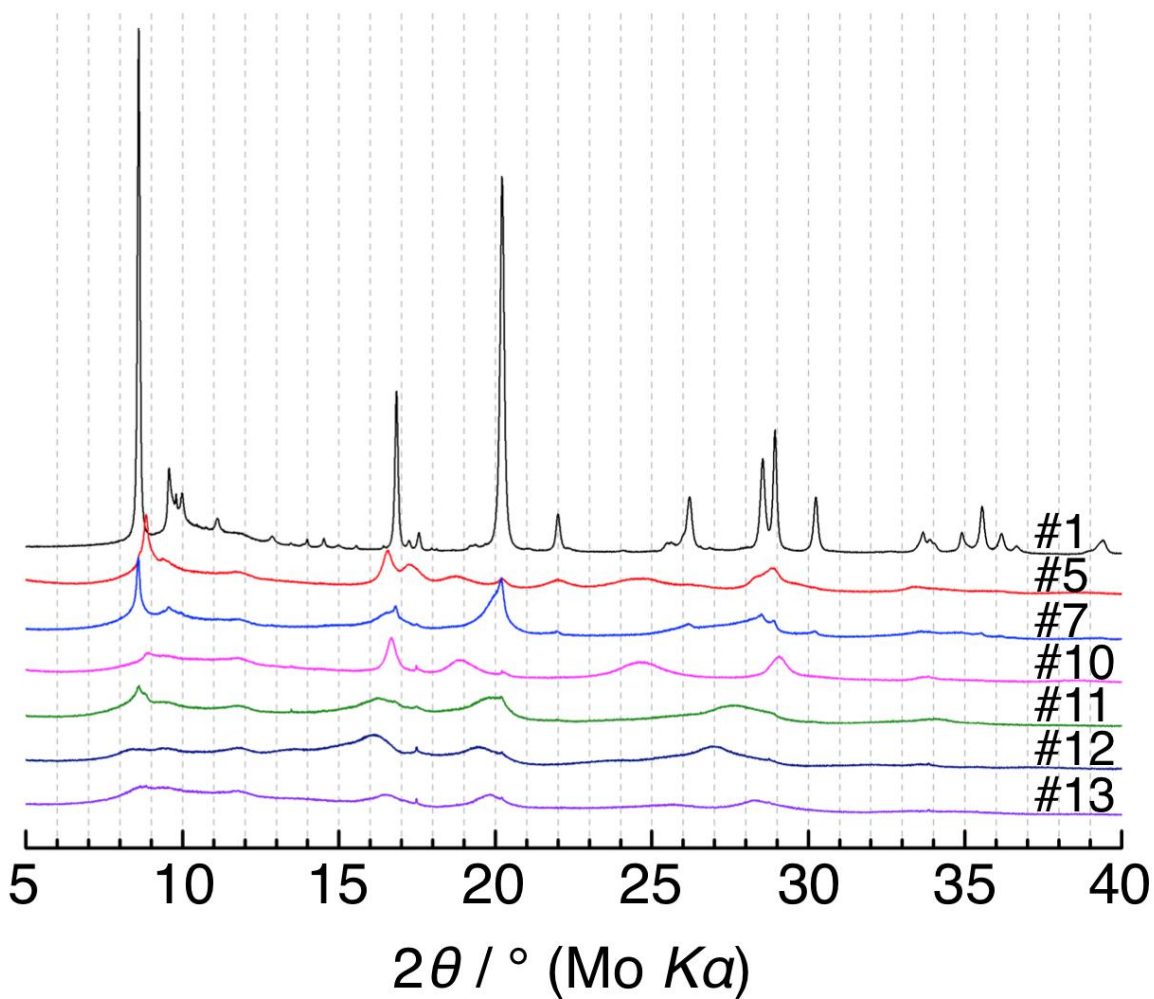


Fig. 7. SR-XRD profiles of the Li_2MnO_3 electrode for the 1st, 2nd cycles, 20th discharge, and 21st charge. The wavelength is converted to that of Mo $K\alpha$ radiation. The sample numbers are described in Fig. 1 and 6.

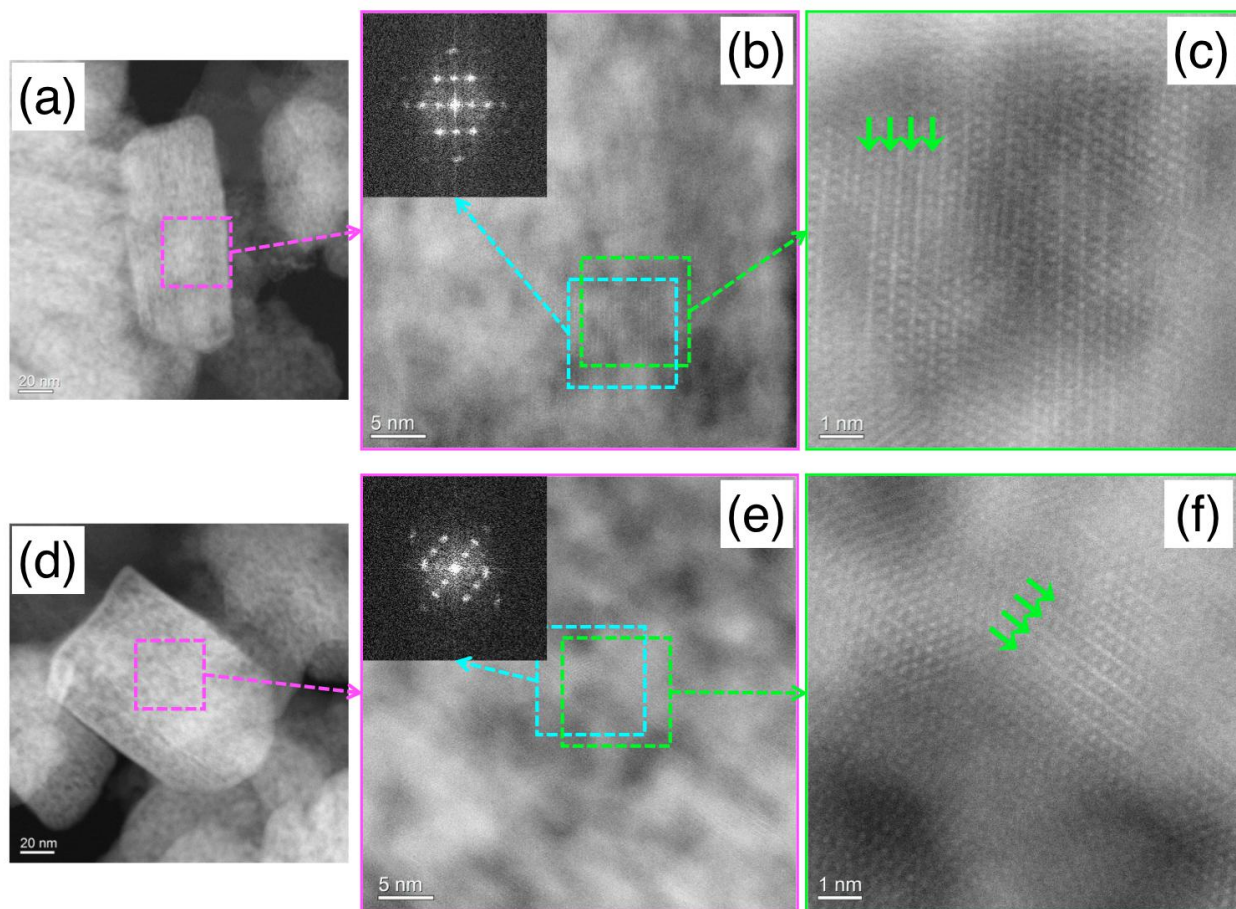


Fig. 8. HAADF-STEM images for the electrode samples disassembled at 2.0 V (upper panels; a,b,c) and 4.8 V (lower panels; d,e,f) in the 20th and 21st cycle, respectively. The FFT patterns of the HAADF images are shown in (b) and (e). The Mn atoms in the Li layers are marked with light green arrows.

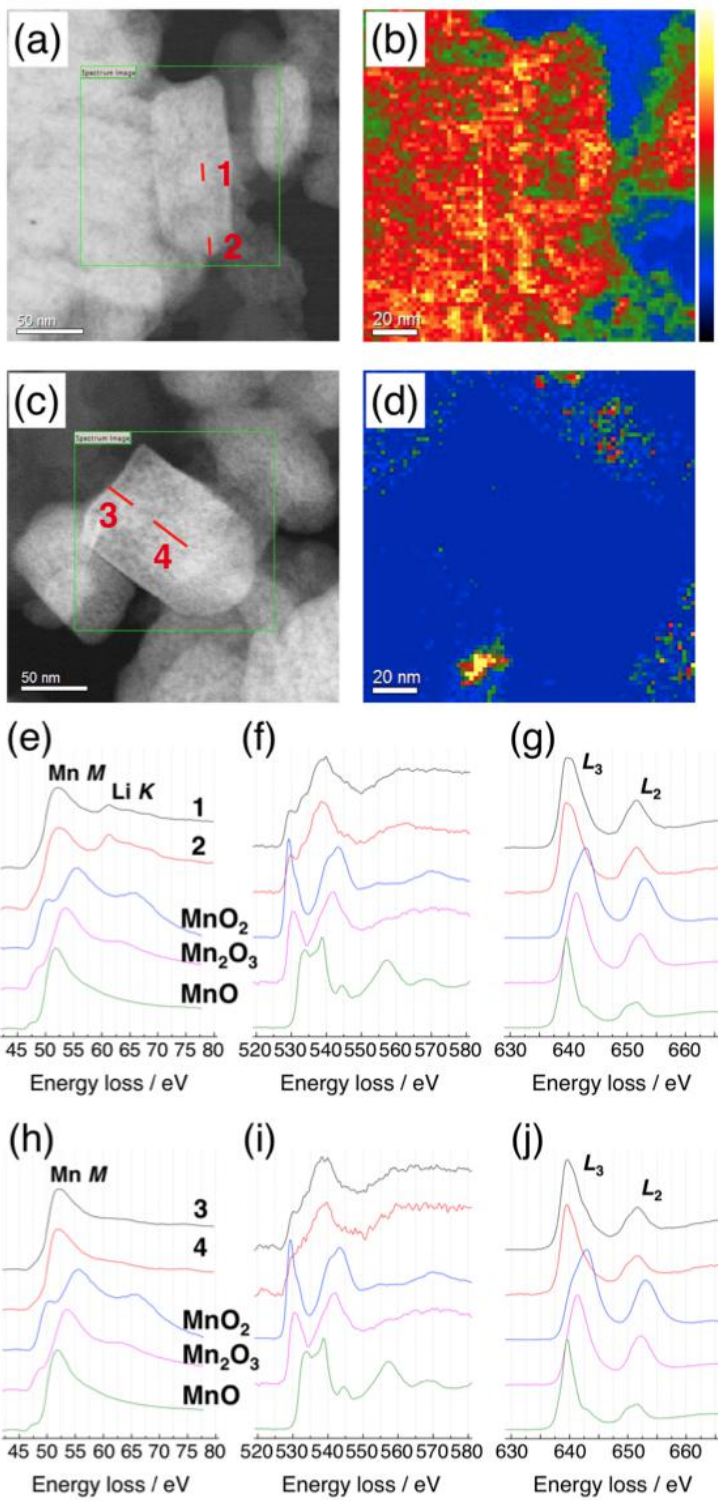


Fig. 9. EELS mapping images and EEL spectra for the electrode samples disassembled at 2.0 V and 4.8 V in the 20th and 21st cycle, respectively. (a) HAADF-STEM image and (b) Li K-edge

map for the sample at 2.0 V. (c) HAADF-STEM image and (d) Li K-edge map for the sample at 4.8 V. The EEL spectra at (e) Mn M-edge + Li K-edge, (f) O K-edge, and (g) Mn L-edge for the sample at 2.0 V, (h) Mn M-edge + Li K-edge, (i) O K-edge, and (j) Mn L-edge for the sample at 4.8 V. The spectra 1–4 are acquired from the lines 1–4 in (a) and (c). The reference spectra of MnO_2 (Mn^{4+}), Mn_2O_3 (Mn^{3+}), and MnO (Mn^{2+}) are also shown.

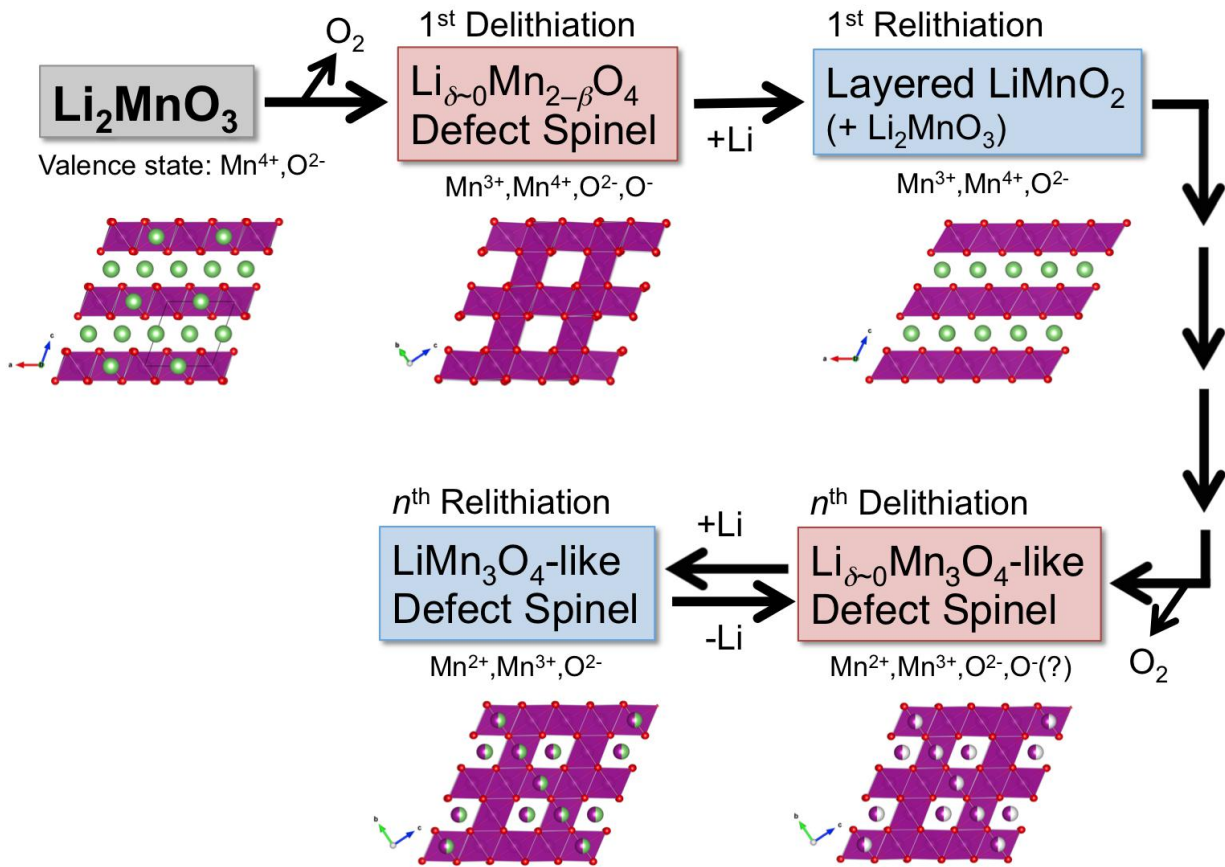


Fig. 10. Schematic illustration of the layered-to-spinel phase transformation in Li_2MnO_3 during the initial cycle and after the multiple cycles.

Table of Contents

Direct Observation of Layered-to-Spinel Phase Transformation in Li_2MnO_3 and Spinel Structure Stabilised after the Activation Process

Keiji Shimoda, Masatsugu Oishi, Toshiyuki Matsunaga, Miwa Murakami, Keisuke Yamanaka, Hajime Arai, Yoshio Ukyo, Yoshiharu Uchimoto, Toshiaki Ohta, Eiichiro Matsubara, Zempachi Ogumi

The layered-to-spinel phase transformation in Li_2MnO_3 during the initial charge occurs in the two-phase reaction process within a single particle.

