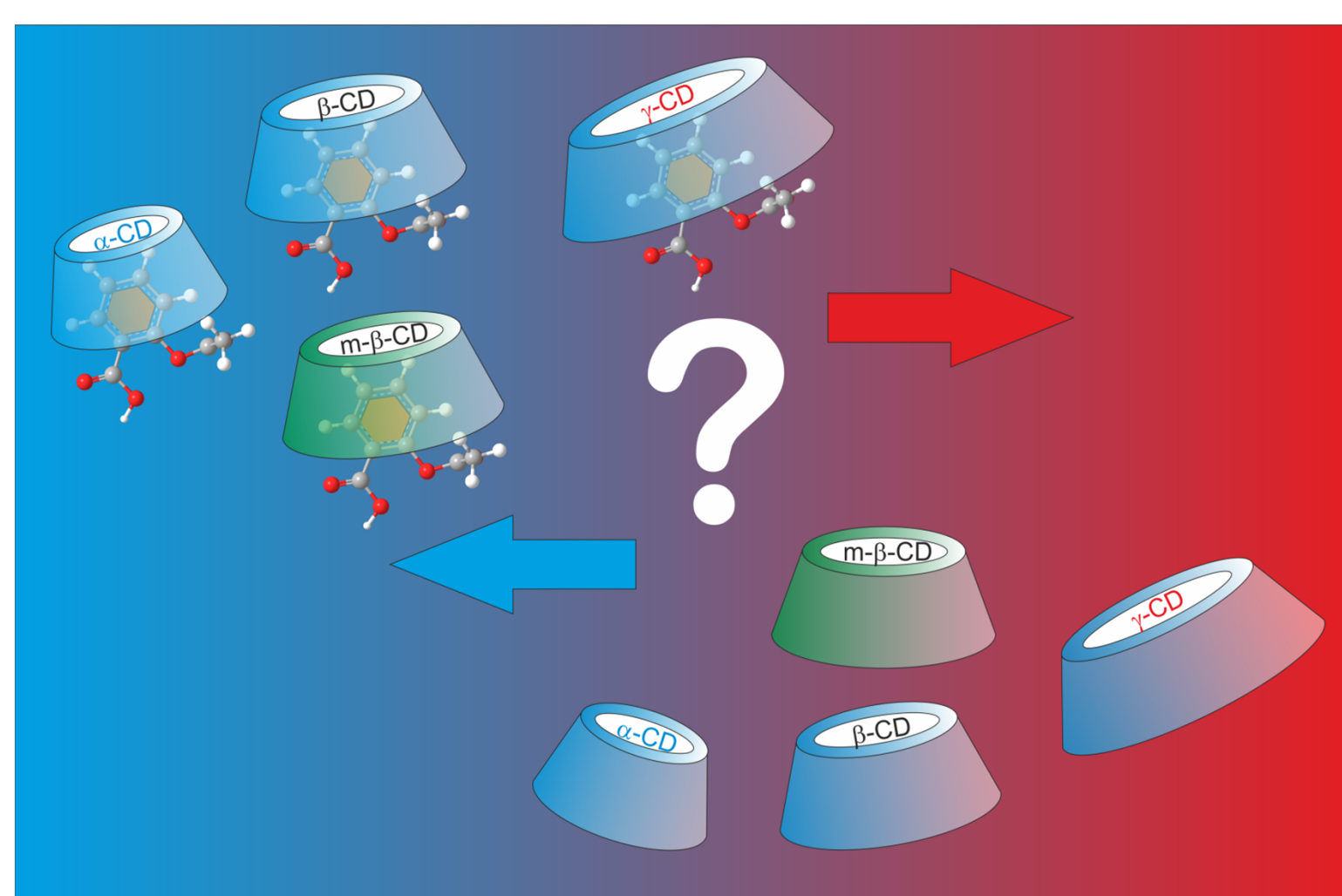
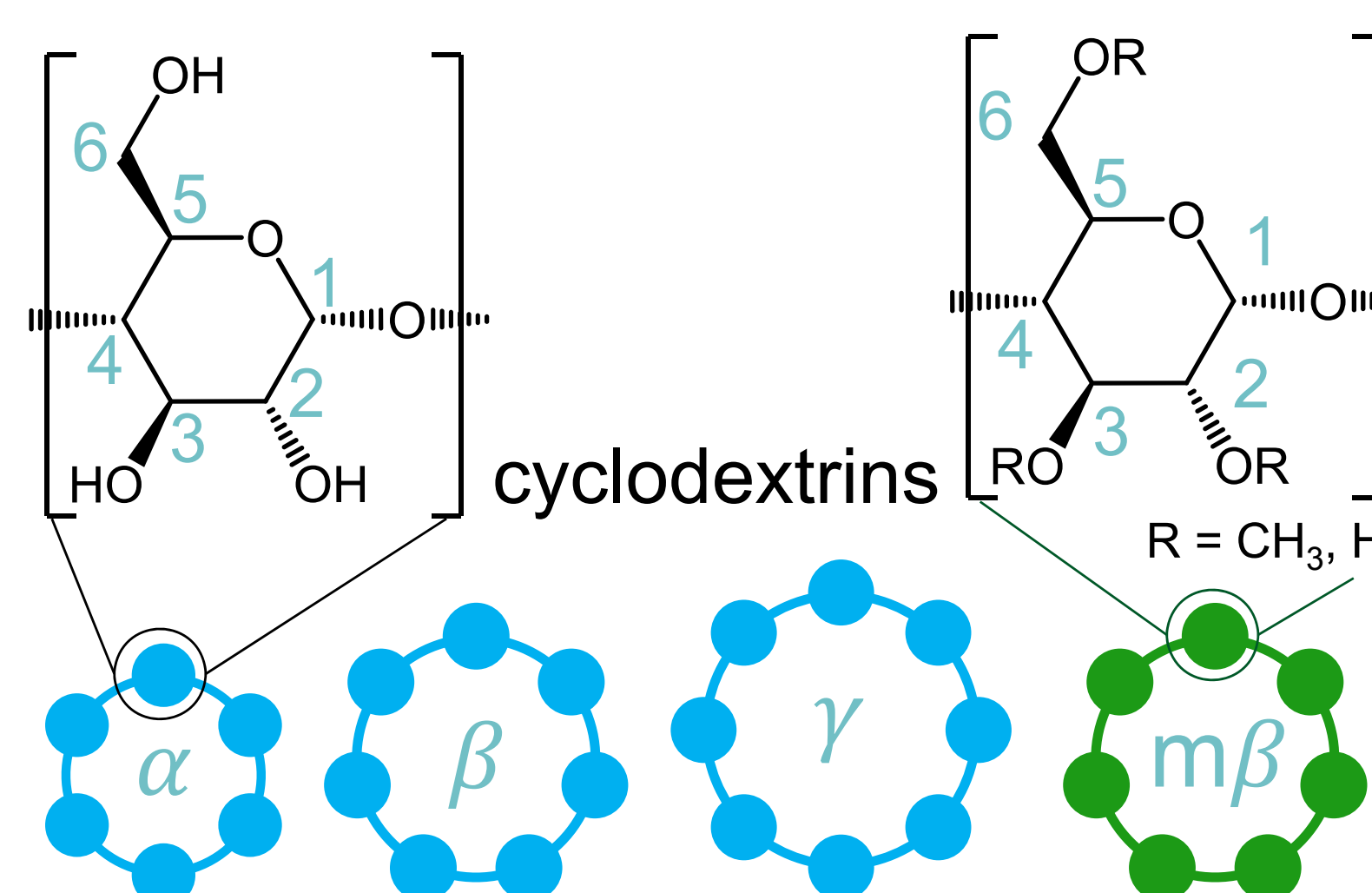


## Motivation [1]

- Controlled movement of drugs into certain, e.g. inflamed, areas
- Check complex formation:  $\alpha$ -,  $\beta$ -,  $\gamma$ - and methyl- $\beta$ -cyclodextrin with aspirin by NMR
- investigation of thermophoretic behavior

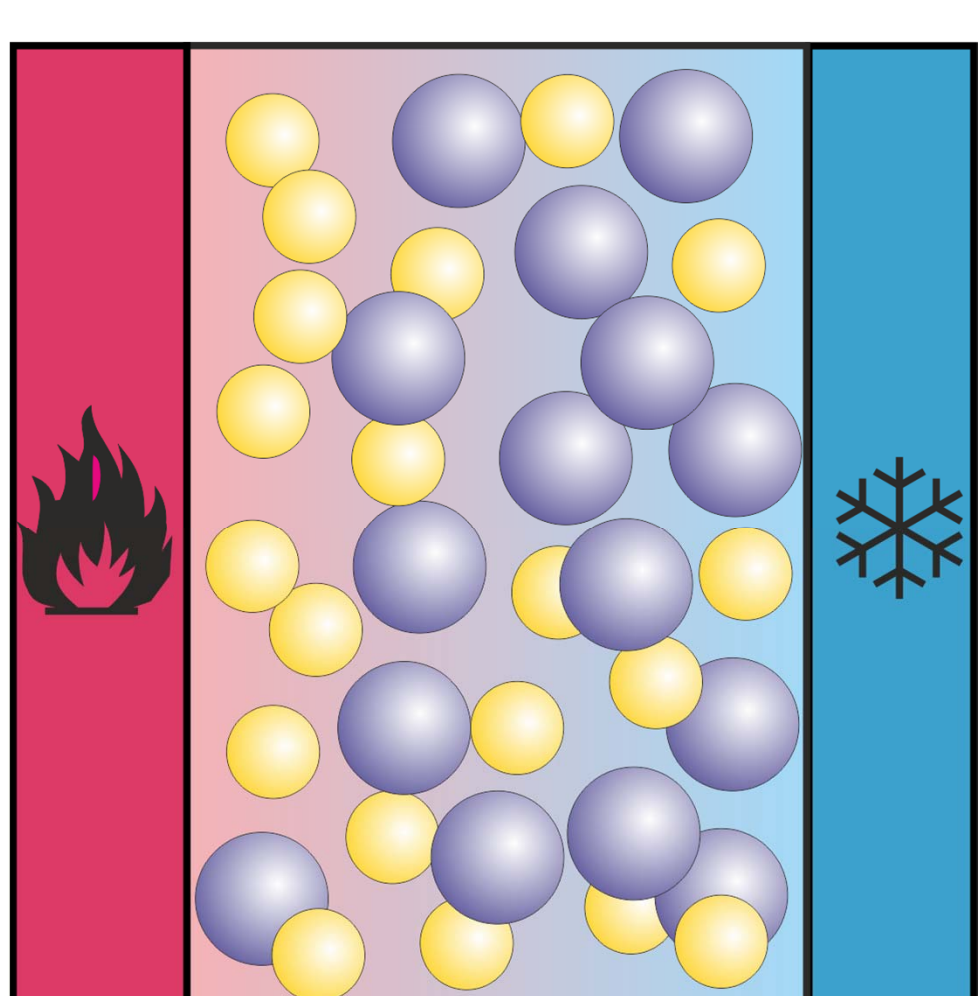


## System



concentration: 1wt% CD  
molar 1:1-ratio: CD:aspirin

## Thermophoresis [2]



- Flux  $\vec{j}$  along a temperature gradient:
- thermal diffusion  $D_T$  along a temperature gradient  $\nabla T$
  - Fickian diffusion  $D$  along the induced concentration gradient  $\nabla c$ .

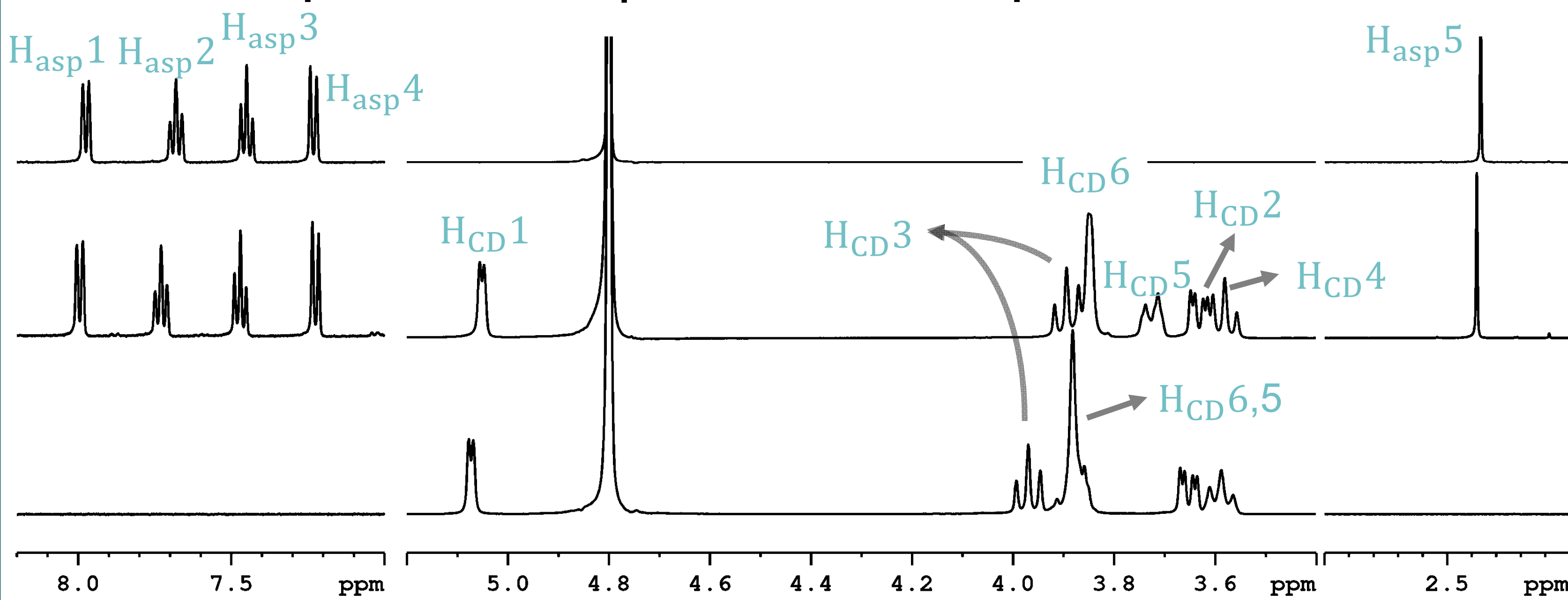
$$\vec{j} = -D\nabla c - c(1-c)D_T\nabla T$$

In the steady state ( $\vec{j}=0$ ) the Soret coefficient  $S_T$  is defined

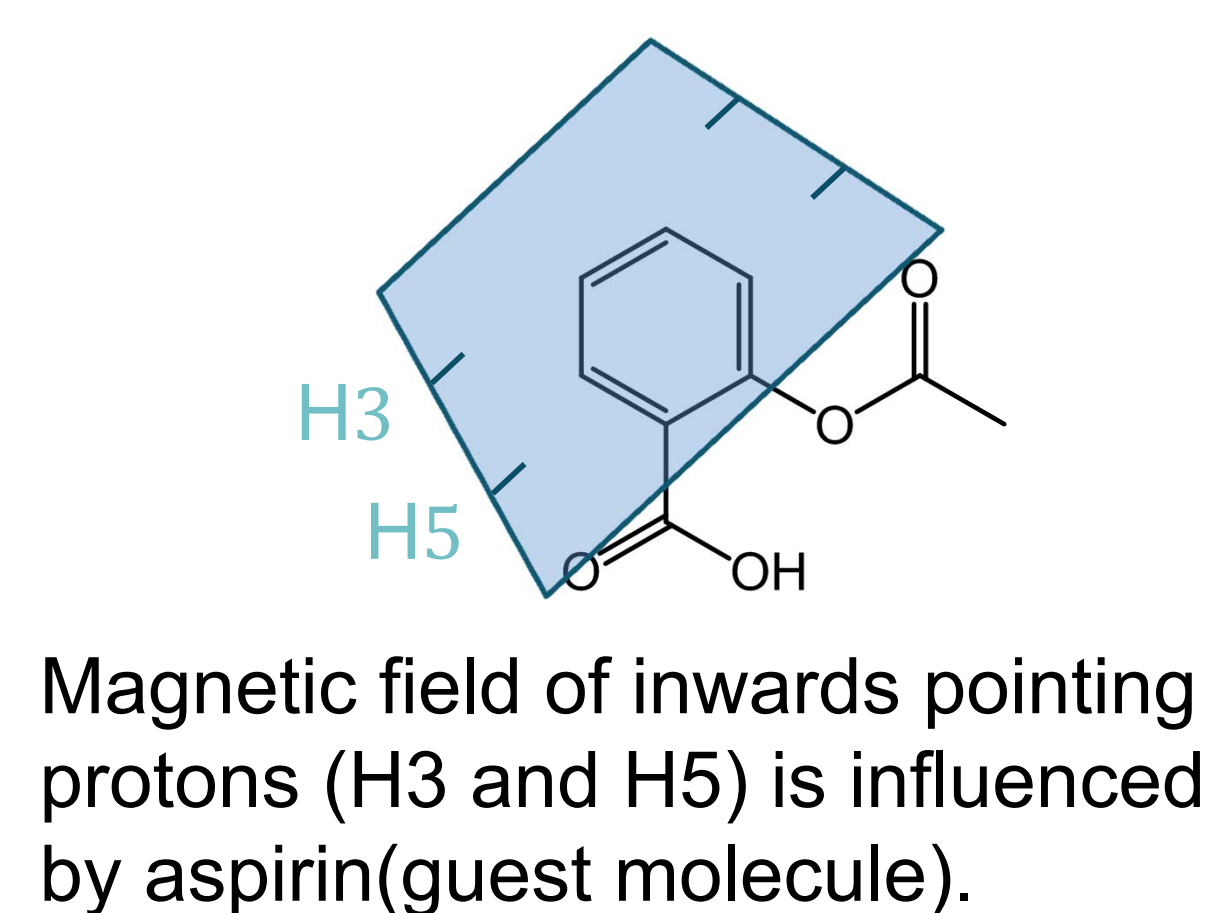
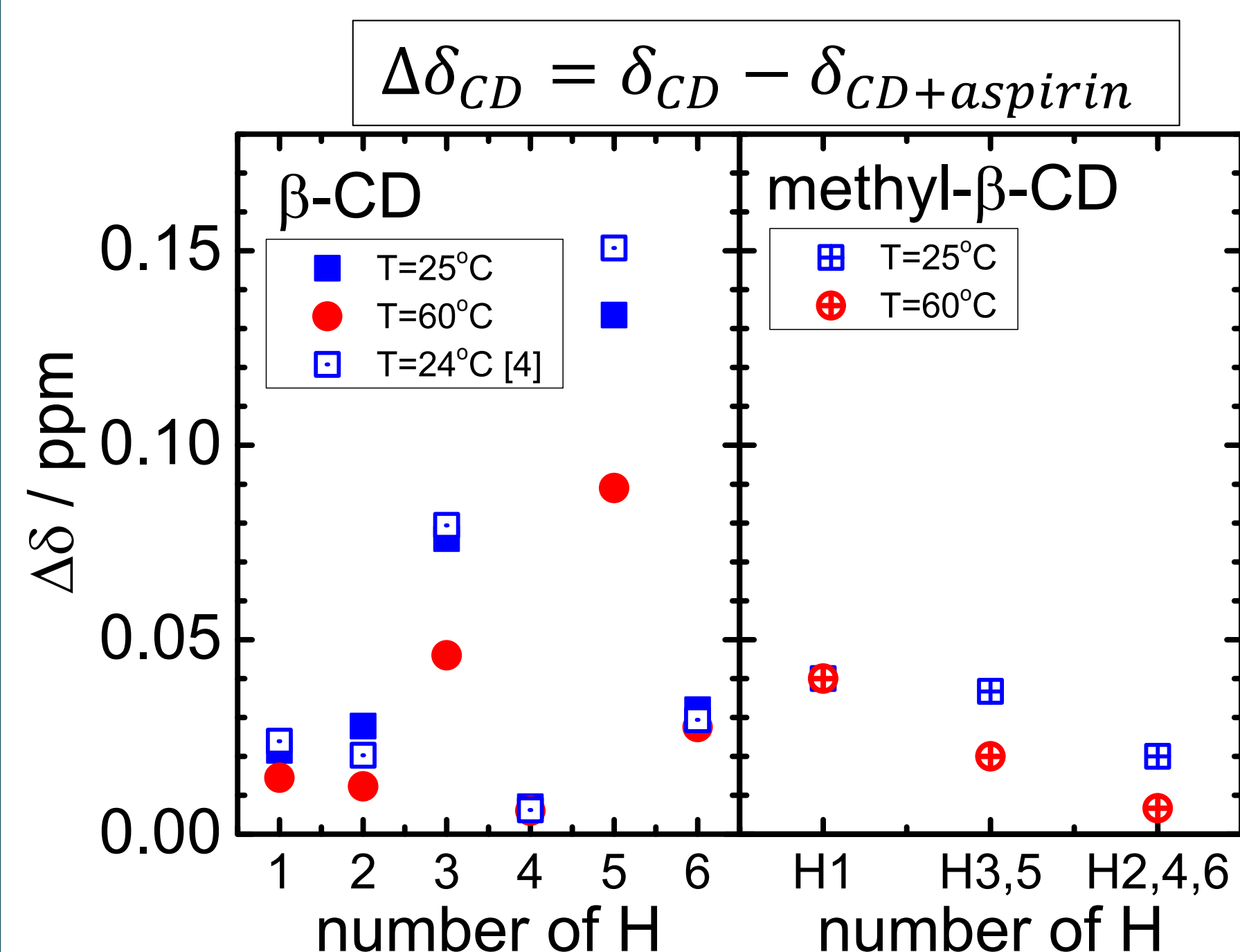
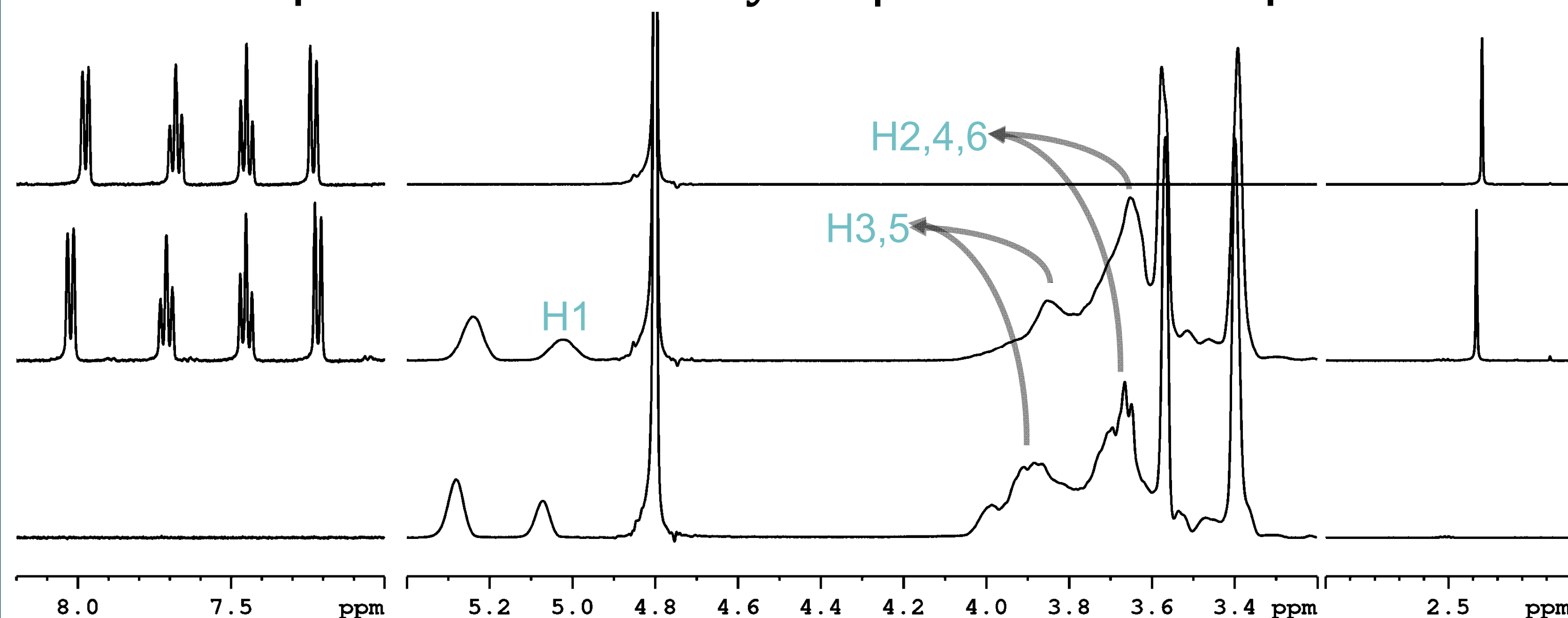
$$S_T \equiv \frac{D_T}{D} = -\frac{1}{c(1-c)} \frac{\Delta c}{\Delta T}$$

## NMR measurements [4], [6]

<sup>1</sup>H NMR spectrum of  $\beta$ -CD and aspirin 25°C



<sup>1</sup>H NMR spectrum of methyl- $\beta$ -CD and aspirin 25°C



$\beta$ -CD

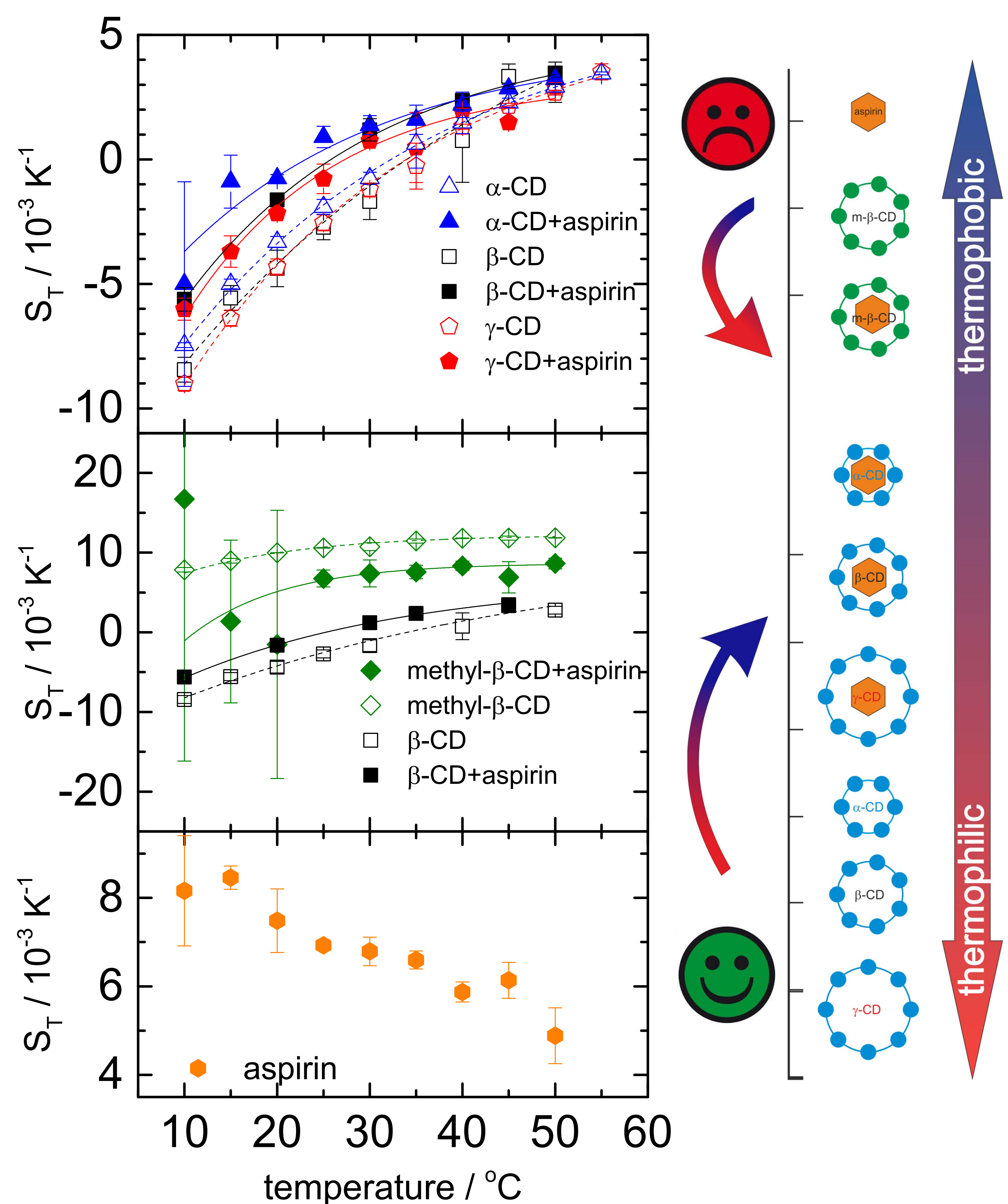
- large  $\Delta\delta$  value is observed for H3 and H5 compared to the other protons.
- $\Delta\delta$  value of H3 and H5 is decreased by heating.

methyl- $\beta$ -CD

- small  $\Delta\delta$  value observed for H3,5 compared to  $\beta$ -CDs H3 and H5.
- $\Delta\delta$  value of H3,5 is decreased by heating.

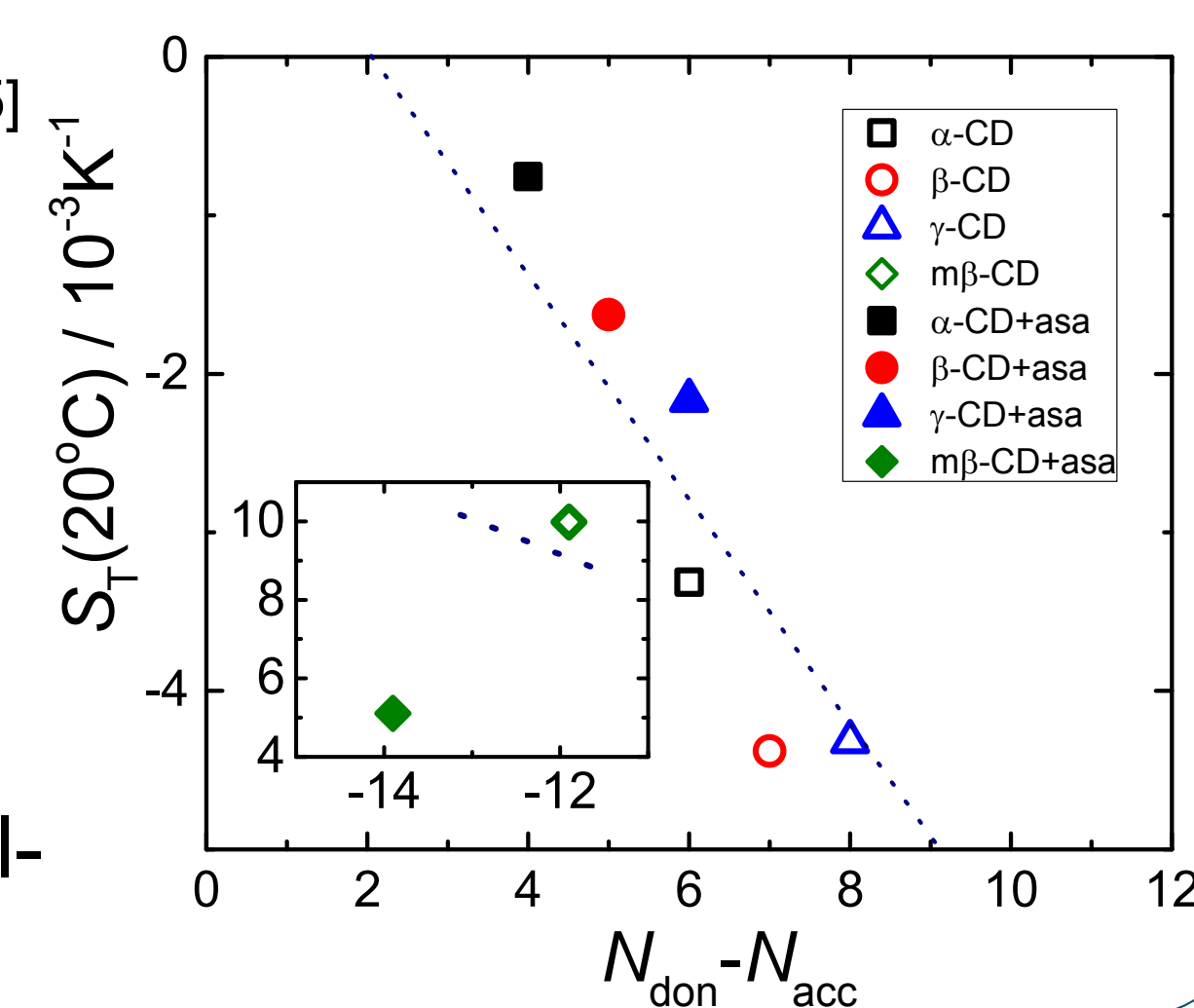
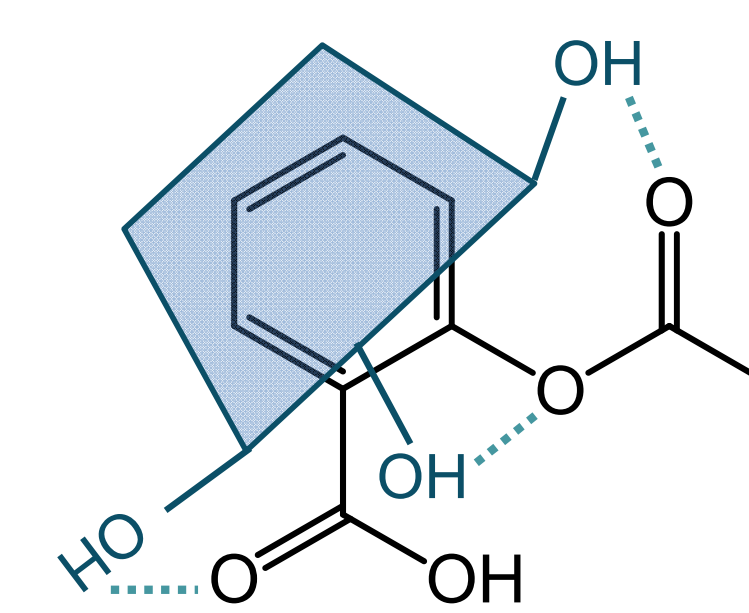
**Inclusion complex is confirmed for both CDs.**

## TDFRS results [3]



## Conclusion

- complexes formation are confirmed between CD and aspirin from NMR measurement
- At the high temperature the complex capability weakens
- behaviour of pure CDs and un-methylated CD-complexes fits donor-acceptor model<sup>[5]</sup>
- fewer HB sites result in a higher  $S_T$ /stronger thermophobicity
- drug-complex of methyl- $\beta$ -CD behaves different
- possible explanations are polarisation or charge effects
- more information about structure of methyl- $\beta$ -CD-complex is needed



## References

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