

Experimental and kinetic study of thermal decomposition behaviour of phytoremediation derived *Pteris vittata*

Authors: Lunbo Duan^{1, 2*}, Jian Chen¹, Ying Jiang², Xiaole Li¹, Philip Longhurst², Mei Lei³

1. Key Laboratory of Energy Thermal Conversion and Control, Ministry of Education, School of Energy and Environment, Southeast University, Nanjing, 210096, China;

2. School of Water, Energy and Environment, Cranfield University, Cranfield, Bedfordshire, MK43 0AL, UK

3. Centre for Environmental Remediation, Institute of Geographic Sciences and Natural Resources Research, Chinese Academy of Sciences (CAS), Beijing, 100101, China

*Corresponding author: Tel: +86(0)25 8379 0147, E-mail: duanlunbo@seu.edu.cn

ABSTRACT

Combustion and gasification for biomass to energy conversion is often suggested for the management of residual *P. vittata* from phytoremediation. In this study, the thermal behaviour of *P. vittata* was studied on a thermogravimetric analyser (TG), and the kinetic triplet of biomass sample was further determined for different stages of the thermochemical processes using the Ozawa and KAS methods, subsequently modified by an iterative procedure. Results show that thermal decomposition under combustion condition was complete at a lower temperature of ~500°C compared to ~700 °C for gasification, indicating the both easily complete conversion of *P. vittata* by combustion and gasification. Kinetic study shows that although activation energy for each stage under combustion condition is mostly larger than that under gasification, yet the reaction rate of thermal decomposition of *P. vittata* under combustion condition is still great larger than that under gasification condition. These findings strongly suggest that thermochemical processes offer suitable methods for the volume reduction and energy production of *P. vittata*.

KEYWORDS: *P. vittata*, Phytoremediation, Waste biomass, TG, Thermochemical conversion, Kinetic analysis

1. Introduction

Globally, arsenic and heavy metals have become the most significant inorganic soil contaminant following intensive mining and associated industrial activities. It is estimated that anthropogenic release of arsenic to soil are 28,400–94,000 tons per year [1]. In China, arsenic pollution is of particularly concern in the mining regions and the vicinity farmland where toxic contaminants are transported from the source points by rainfall or floodwater [2]. The affected regions cover extensive areas and are often in economic deprived rural districts, therefore a cost-effective alternative to conventional remediation approach is required for the treatment and pollution management of these sites. Compared with conventional land remediation

strategies for heavy metal contaminations, phytoremediation using plants to remove the contaminants from the soil to plant biomass offers significant economic, aesthetic, and technical advantages [3, 4].

The success of phytoremediation projects depends on a variety of factors, of which phytoremediation plant selection is key. For phytoremediation of arsenic in particular, there are numerous arsenic hyperaccumulators that have been reported to date, among which the majority belong to the fern *Pteris* family [5-10] and one particular species, *P. vittata* has been proven successful in many field studies of arsenic phytoremediation [11-13].

However, a key inhibiting factor for a wider implementation of *P. vittata* is the disposal of large quantities of arsenic contaminated plant biomass material throughout the process [14, 15]. It is reported that arsenic concentrations in *P. vittata* can accumulate up to several thousand mg kg⁻¹ [16], which causes secondary environmental pollution if not properly treated. Thermochemical biomass to energy conversion technologies, i.e. combustion, pyrolysis and gasification, has been proposed as a feasible method to achieve significant volume reduction of phytoremediation derived biomass and bioenergy production [17, 18].

To date, there is little literature record that provides detailed information on the thermal behaviour of phytoremediation biomass, particularly on *P. vittata*. Phytoremediation biomass thermal behaviour study is essential to provide key operational parameters for thermochemical processes optimisation, and further the understanding of bioenergy potential of this particular type of residual biomass. In the development of using phytoremediation as a viable land remediation method and converting biomass residuals from remediation as a bioenergy sources, understanding of the thermal behaviour and reactivity of *P. vittata* is of paramount importance.

Thermogravimetric analysis (TG) is a widely applied technique to study the primary reaction in the decomposition and the thermal behaviours of biomass. TG results, when appropriately interpreted can provide key information of thermochemical reaction stages and kinetics of biomass conversion under a specific thermochemical condition [19], which is particularly important for the burner design and products quality optimisation.

There are a number of mathematical methods proposed to calculate the kinetic parameters based on TG results. These methods can be broadly categorised into model-fitting and iso-conversional methods [20]. Model fitting methods involve fitting different models to temperature curves and simultaneously determining the activation energy E and pre-exponential factor A . This normally requires significant efforts of trial-and-error fitting of the curve and can potentially lead to false results, if an inappropriate kinetic model is selected.

Iso-conversional or "model-free" calculation methods calculate activation energy at progressive degrees of conversion α without modelistic assumptions. Iso-conversional methods including Kissinger-Akahira-Sunose (KAS) method [21] and Ozawa method [22, 23] are well established and widely accepted due to their simplicity and avoidance of errors resulted from modelistic assumptions [24]. The errors and inaccuracy associated with these iso-conversional methods are well understood [25, 26]. During the past decade, an iterative procedure is proposed and often used by researchers to improve the accuracy of results produced from conventional iso-conversional methods [25].

In this study, thermogravimetric analysis (TG) was conducted to study the thermal behaviour of *P. vittata*, and the kinetic triplet of the biomass sample was further determined. To the best

of the authors' knowledge, this is the first time that the kinetic triplet of *P. vittata* harvested from a phytoremediation field trial has been reported. The thermal behaviour and kinetic information reported in this study will provide valuable information for the follow up experiments using fixed bed and fluidised bed facilities and future scale up thermochemical treatment of phytoremediation residual biomass.

2. Materials and Methods

2.1 Sample Preparation

P. vittata was harvested from a large scale arsenic phytoremediation field trial in South Central China, where the studied site was contaminated with high concentration of arsenic due to flood runoff from nearby mining sites. The arsenic concentrations in biomass and growth soil were determined using Inductively Coupled Plasma - Mass Spectrometry (ICP-MS) following acid digestion [27] and are 237 and 25.3 mg kg⁻¹, respectively. The proximate analysis of the *P. vittata* was performed according to ASTM E871 (moisture content), ASTM D1102 (ash content) and ASTM E872 (volatile) standards. Fixed carbon content was calculated by difference. The ultimate analysis was conducted according to ASTM E777 (carbon, hydrogen), ASTM E778 (nitrogen) and ASTM E775 (sulphur) standards. Oxygen content was also determined by difference. ASTM D2015, E711 standard was used to determine the heating value [28, 29]. All the above characteristics of the *P. vittata* are listed in Table 1.

Prior to the thermogravimetric study, *P. vittata* was washed with tap water to remove possible adhering soil and dust, followed by rinsing with deionized water. The sample was then dried at 70 °C in an oven for 48 hours [30]. The dried samples were pulverized by agate mortar and sieved through an 80 mesh sieve to obtain particle sizes of less than 0.2 mm.

Table 1 Proximate and ultimate analysis of *P. vittata*

Proximate analysis /wt% (air dry basis)	
Moisture	9.66
Volatiles	57.46
Fixed carbon	17.38
Ash	15.50
Ultimate analysis /wt% (air dry basis)	
Carbon	39.04
Hydrogen	4.38
Nitrogen	1.41
Sulphur	0.32
Oxygen (by difference)	29.69
Heating value /MJkg⁻¹ (dry basis)	
Higher heating value	18.17
Lower heating value	16.93

2.2 Equipment and Method for TG Studies

A thermogravimetric analyser (Perkin-Elmer Pyris 1, USA) was used to investigate the thermal behaviour of *P. vittata* under combustion and gasification conditions. The combustion tests were carried out using air as carrier gas, whereas CO₂ was used for gasification tests. Under both test conditions, the carrier gas flow rate was set at 100 ml min⁻¹. Before starting each test,

the analyser was purged with carrier gas for a minimum of 1 hour. Approximately 3 mg of the biomass sample was used for each run in this study. Under both combustion and gasification conditions, temperature was raised from room temperature to 800 °C with a heating rate of 20, 30, 50 and 80 °C min⁻¹. Each experiment run was conducted in duplicate for data quality assurance.

2.3 Kinetics Study

The fundamental isothermal rate equation applied in all kinetic studies in this work is expressed as Eqs. 1 and 2 [31]:

$$\frac{d\alpha}{dT} = k \cdot f(\alpha) = A \cdot f(\alpha) \cdot \exp(-E/RT) \quad (1)$$

$$\alpha = \frac{\omega_i - \omega_t}{\omega_i - \omega_f} \quad (2)$$

where α is the degree of conversion in %; $f(\alpha)$ is the conversion function; k is the rate constant of reaction, ω_i is the initial mass of sample; ω_t is the mass of sample at time t ; ω_f is the final mass of sample; R is the gas constant at 8.314 J K⁻¹ mol⁻¹; T is the absolute temperature in Kelvin; A is the pre-exponential factor in min⁻¹, and E is the activation energy in kJ mol⁻¹. A , E and $f(\alpha)$ are referred to as kinetic triplet. The specific form of conversion function $f(\alpha)$ relies on the actual reaction mechanism and the algebraic expressions of functions commonly applied in solid-state reactions are well documented elsewhere [32].

For non-isothermal thermogravimetric analysis at constant heating rate β (K min⁻¹) = dT/dt , Eq. 1 can be modified into a non-isothermal rate expression Eq. 3 [33]:

$$\frac{d\alpha}{dT} = \frac{A}{\beta} \cdot f(\alpha) \cdot \exp(-E/RT) \quad (3)$$

The integral of Eq. 3 can be rearrangement as follows into Eq. 4 and subsequently the integration variable can be redefined as $x = E/RT$:

$$g(\alpha) = \int_0^\alpha \frac{d\alpha}{f(\alpha)} = \int_0^T A \cdot \exp(-E/RT) dT \quad (4)$$

$$\xrightarrow{x=E/RT} \frac{AE}{\beta R} \int_x^\infty \frac{\exp(-x)}{x^2} dx = \frac{AE}{\beta R} p(x)$$

where $g(\alpha)$ is the integral form of the conversion function $f(\alpha)$ and $p(x)$ is the temperature

integral presented as $p(x) = \int_x^\infty \frac{\exp(-x)}{x^2} dx$. Owing to the fact that $p(x)$ function can't be

obtained directly via analytical approaches, to date substantial researches have been carried out to develop computational methods for approximate solutions. Herein the values of activation energy were determined using conventional iso-conversional methods, i.e. Ozawa method (Eq. 5) and KAS method (Eq. 6) [34].

$$\ln \beta = \ln \frac{0.0048AE}{Rg(\alpha)} - 1.0516 \frac{E}{RT} \quad (5)$$

$$\ln \frac{\beta}{T^2} = \ln \frac{AE}{Rg(\alpha)} - \frac{E}{RT} \quad (6)$$

In Ozawa and KAS methods, linear least-squares method can be applied to plot $\ln\beta$ versus T^{-1} and the plots of $\ln(\beta T^{-2})$ versus T^{-1} at various degrees of conversion α (between 0.1- 0.9, with 0.1 increment was taken in this study), respectively. The slopes of these two plots are therefore used to estimate the values of activation energy derived from Ozawa and KAS methods. However, due to the application of the approximate value of $p(x)$, errors have been proved to occur during the determination of activation energy [25, 26]. In order to obtain more accurate results closer to 'exact' values of activation energy, an iterative iso-conversional procedure for Ozawa and KAS equations was applied to determine activation energy according to the following equations [35-38]:

$$\ln \frac{\beta}{H(x)} = \ln \frac{0.0048AE}{Rg(\alpha)} - 1.0516 \frac{E}{RT} \quad (7)$$

$$\ln \frac{\beta}{h(x)T^2} = \ln \frac{AE}{Rg(\alpha)} - \frac{E}{RT} \quad (8)$$

where $h(x)$ and $H(x)$ are expressed as follows in Eqs. 9 and 10:

$$h(x) = \frac{x^4 + 18x^3 + 88x^2 + 96x}{x^4 + 20x^3 + 120x^2 + 240x + 120} \quad (9)$$

$$H(x) = \frac{\exp(-x)h(x)/x^2}{0.0048 \exp(-1.0516x)} \quad (10)$$

The iterative procedure performed in this paper is described in detail as follows [25].

First step: Assuming $h(x)=1$ and $H(x)=1$ to evaluate the initial value of the activation energy E_1 . The conventional iso-conversional methods finish the calculation at this step.

Second step: According to the value of activation energy E_1 , $h(x)$ and $H(x)$ are obtained, and then a new value of activation energy E_2 is calculated from the plot of $\ln\beta \cdot H^{-1}(x)$ versus T^{-1} or the plot of $\ln(\beta (h(x) T^2)^{-1})$ versus T^{-1} on basis of Eq. 7 or 8.

Third step: Repeating Second step and replacing E_1 with E_2 . The procedure is not finished until the absolute difference of $(E_i - E_{i-1})$ is less than a defined small quantity. The last value E_i is the exact value of activation energy. Additionally, herein the 'defined small quantity' (ΔE) used during the iterative calculations was 0.1 kJ mol⁻¹.

The error of activation energy obtained from conventional iso-conversional methods, i.e. Ozawa method and KAS method, without performing the iterative procedure is calculated by Eq. 11.

$$\text{Error}(\%) = 100 \cdot \frac{|E_1 - E_i|}{E_i} \quad (11)$$

where E_1 is the initial value of the activation energy when assuming $h(x) = 1$ and / or $H(x) = 1$ in conventional iso-conversional methods. E_i is the exact value of activation energy.

Kinetic parameters are strongly dependent on the selection of mechanism functions, herein the degrees of conversion corresponding to multiple heating rates taken from the same

temperature were put into 35 types of commonly used mechanism functions which is well documented elsewhere [32]. Linear least-square method can be applied to plot of $\ln\beta$ versus $\ln g(\alpha)$ defined in Eq. 12, therefore determining the most probable mechanism function [39].

$$\ln g(\alpha) = \left[\ln \frac{AE}{R} + \ln \frac{\exp(-x)}{x^2} + \ln h(x) \right] - \ln \beta \quad (12)$$

Since the $g(\alpha)$ function is determined, the value of pre-exponential factor can be calculated from the intercept of the plots of Eq. 6.

3. RESULTS AND DISCUSSION

3.1 Influence of thermal treatment on decomposition of *P. vittata*

Thermogravimetric (TG) curves in Fig. 1 clearly show that mass of sample starts to decrease immediately as temperature raised. Although thermal decomposition under combustion is complete at a lower temperature of $\sim 500^\circ\text{C}$ compared to $\sim 700^\circ\text{C}$ for gasification, there is no significant difference in final mass between combustion (7.9% left) and gasification (8.1% left) conditions. This phenomenon may be explained that there was almost no moisture, volatile matter and fixed carbon content existed in the sample and only ash content was remaining after thermal treatments (combustion or gasification). The result suggests that both combustion and gasification can easily realize complete conversion of *P. vittata*.

Differential thermogravimetric (DTG) curves in Fig. 1 indicate that similar to other thermal decomposition study on biomass, there are three distinct mass loss stages during decomposition of *P. vittata* sample, i.e. dehydration, pyrolysis and char conversion [20]. Dehydration stage occurs due to the loss of water present in the sample and external water bounded by the surface tension [40]. For both combustion and gasification conditions, the temperature range for dehydration is between ~ 30 - 180°C , similar to the results observed in previous studies on other biomass samples [41, 42]. Pyrolysis stage of biomass thermal decomposition is caused by volatilisation of hemicellulose, cellulose and part of the lignin contents in the biomass [43, 44]. Temperature ranges for pyrolysis stage are very approached under combustion and gasification conditions. The temperature ranges under combustion and gasification conditions are between ~ 180 - 405°C and ~ 180 - 425°C , respectively. The slightly lower temperature range during pyrolysis stage under combustion indicates the presence of oxygen promotes the decomposition of biomass sample [45]. However, char conversion stage of biomass thermal decomposition differs a lot when different thermal treatments are selected. Although char conversion stage under both conditions is partly assigned to the decomposition of the remaining lignin, yet char conversion stage under combustion is also related to the combustion of char residue [46] while under gasification is char gasification [47, 48]. The maximum mass loss rate for this stage under combustion condition is $-0.5\%^\circ\text{C}^{-1}$, which is great larger than that under gasification, as a value of $-0.16\%^\circ\text{C}^{-1}$. Meanwhile, the temperature ranges under combustion and gasification conditions are between ~ 405 - 500°C and ~ 425 - 700°C , respectively. Both the greatly lower temperature range and significantly larger maximum mass loss rate during char conversion stage under combustion indicate the presence of oxygen promote the decomposition of biomass sample as well.

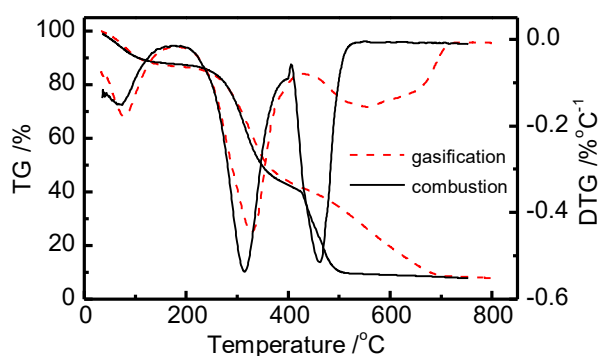


Fig. 1 TG and DTG curves for *P. vittata* in combustion and gasification processes under the heating rate of 50 °C min⁻¹

3.2 Influence of heating rate during thermal treatment of *P. vittata*

Figs. 2 and 3 show the TG curves with a range of heating rates under combustion and gasification conditions, respectively. As shown in Fig. 2 or 3, most of the curves exhibit the same variation tendency under combustion or gasification condition with the changes of heating rate. Additionally, changing heating rate doesn't lead to any change of final mass under both conditions. The results indicate that heating rate has a negligible effect on final mass of *P. vittata* under thermal treatment conditions, which is coherent with the conclusion obtained by Nie et al. [49].

Figs. 4 and 5 show the DTG curves with a range of heating rates under combustion and gasification conditions, respectively. Just like DTG curves presented in Fig. 1 (under combustion and gasification condition with a heating rate of 50 °C min⁻¹), all the DTG curves at various heating rates can be still separated into three distinct mass loss stages under combustion or gasification condition. Meanwhile, it is interesting to find that heating rate has negligible effect on the previous two stages while significant effect on the char conversion stage. DTG curves of char conversion stage under combustion and gasification conditions exhibit dramatically different variation trend with the change of heating rate. For char conversion stage under combustion condition, an increase of heating rate leads to the shift of DTG curve towards higher temperature and an enhancement in the mass loss peak of sample in DTG curve as a result of the heat transfer limitations [41, 50]. While for char conversion stage under gasification, increasing heating rates leads to an enhancement in the mass loss peak of sample except for the curve of heating rate of 20 °C min⁻¹ and doesn't cause any shift of DTG curves towards higher temperature, which still need to be further investigated. Despite of the atmosphere, the third stage both consists of two processes, i.e. the decomposition of remaining lignin and the further conversion of generated char residue. Since the lignin decomposition should be very similar, it can be inferred that the differences were caused by the combustion and gasification of char.

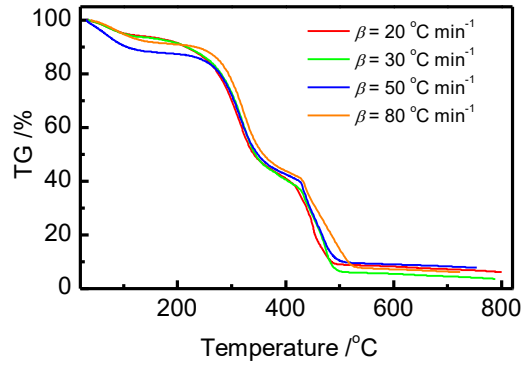


Fig. 2 TG curves of *P. vittata* recorded under combustion condition at four heating rates

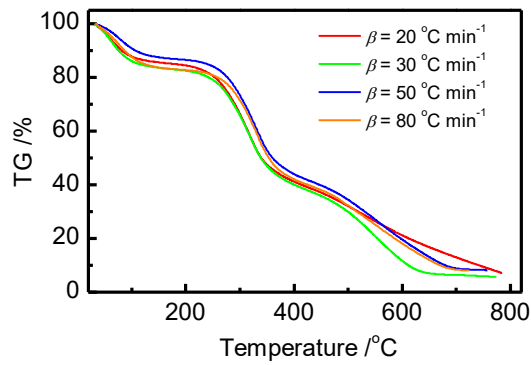


Fig. 3 TG curves of *P. vittata* recorded under gasification condition at four heating rates

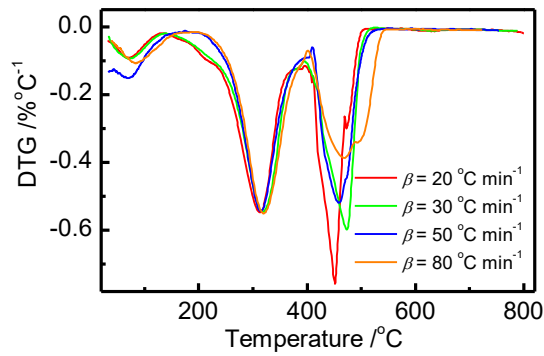


Fig. 4 DTG curves of *P. vittata* recorded under combustion condition at four heating rates

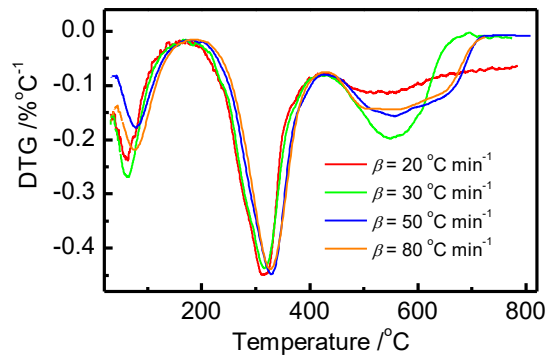


Fig. 5 DTG curves of *P. vittata* recorded under gasification condition at four heating rates

3.3 Kinetic analysis

Kinetic analysis was done to provide key information of thermochemical reaction stages and kinetics of biomass conversion. Since there are three distinct mass loss stages for combustion and gasification and the formation mechanisms for each stage differ greatly, the determinations of kinetic triplet were conducted for each stage instead of whole stage, aiming to obtain more precise parameters. The values of activation energy for each stage determined by Ozawa and KAS method, as well as the error of activation energy obtained from conventional iso-conversional methods without performing the iterative procedure, are tabulated in Tables 2 and 3, respectively.

Table 2 Activation energy E determined from conventional iso-conversional methods (Ozawa method, KAS method) and corresponding iterative iso-conversional procedure for each stage under combustion condition

	α	Ozawa method /kJmol ⁻¹	Iterative of Ozawa /kJmol ⁻¹	Error /%	KAS method /kJmol ⁻¹	Iterative of KAS /kJmol ⁻¹	Error /%
Stage I dehydration	0.1	121.39	122.41	0.83	122.31	122.41	0.75
	0.2	98.40	97.00	1.44	95.91	97.00	0.31
	0.3	89.90	85.33	5.36	84.61	85.33	4.53
	0.4	76.41	76.82	0.53	74.89	76.82	1.99
	0.5	72.66	70.25	3.42	68.72	70.25	1.24
	0.6	67.92	65.80	3.23	64.79	65.79	1.70
	0.7	63.41	61.00	3.96	59.56	61.00	1.60
	0.8	57.39	55.17	4.03	52.71	55.16	0.42
	0.9	45.35	42.50	6.70	40.27	42.50	1.45
	average	76.98	75.14		73.75	75.14	
Stage II pyrolysis	0.1	64.10	60.01	6.82	58.41	60.01	2.65
	0.2	112.32	109.35	2.72	109.04	109.35	0.29
	0.3	159.30	158.41	0.57	158.16	158.40	0.15
	0.4	207.05	203.96	1.51	203.79	203.96	0.08
	0.5	236.08	238.60	1.06	239.15	238.60	0.23
	0.6	259.02	262.56	1.35	262.42	262.56	0.05
	0.7	267.21	275.85	3.13	276.68	275.85	0.30
	0.8	284.83	284.40	0.15	281.53	284.40	1.01
	0.9	337.78	344.88	2.06	342.11	344.88	0.80
	average	214.19	215.33		214.59	215.33	
Stage III char conversion	0.1	312.71	318.91	1.94	320.11	318.91	0.38
	0.2	344.91	350.61	1.63	354.06	350.61	0.98
	0.3	354.04	364.30	2.82	364.09	364.30	0.06
	0.4	322.50	332.34	2.96	331.54	332.34	0.24
	0.5	275.09	278.18	1.11	283.63	278.18	1.96
	0.6	225.41	226.78	0.60	227.66	226.78	0.39
	0.7	189.89	186.03	2.07	182.66	186.03	1.81
	0.8	181.66	178.84	1.58	177.73	178.83	0.62
	0.9	177.57	174.37	1.84	173.97	174.37	0.23

average	264.87	267.82	268.38	267.82
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Table 3 Activation energy E determined from conventional iso-conversional methods (Ozawa method, KAS method) and iterative iso-conversional procedure for each stage under gasification condition

	α	Ozawa method /kJmol ⁻¹	Iterative of Ozawa /kJmol ⁻¹	Error /%	KAS method /kJmol ⁻¹	Iterative of KAS /kJmol ⁻¹	Error /%
Stage I dehydration	0.1	86.16	87.44	1.47	85.79	87.44	1.89
	0.2	67.59	64.70	4.46	63.04	64.70	2.57
	0.3	57.53	54.99	4.62	54.23	54.99	1.39
	0.4	54.44	50.52	7.77	49.68	50.52	1.65
	0.5	50.95	47.89	6.40	46.89	47.89	2.08
	0.6	48.07	45.78	5.00	44.00	45.77	3.87
	0.7	45.34	43.11	5.17	42.76	43.11	0.81
	0.8	43.19	40.46	6.75	38.33	40.46	5.27
	0.9	36.55	33.74	8.32	32.81	33.74	2.76
average		54.42	52.07		50.84	52.07	
Stage II pyrolysis	0.1	88.43	88.65	0.25	84.59	88.65	4.58
	0.2	121.93	122.04	0.09	118.40	122.04	2.98
	0.3	133.40	134.53	0.84	130.72	134.53	2.83
	0.4	139.43	144.92	3.79	143.71	144.92	0.83
	0.5	151.68	155.50	2.46	154.78	155.50	0.47
	0.6	160.10	162.71	1.60	162.13	162.71	0.36
	0.7	173.03	172.06	0.56	171.18	172.05	0.51
	0.8	178.67	175.39	1.87	176.55	175.39	0.66
	0.9	151.87	151.22	0.43	152.55	151.22	0.88
average		144.28	145.22		143.85	145.22	
Stage III char conversion	0.1	146.42	147.86	0.97	146.43	147.86	0.97
	0.2	185.37	184.95	0.23	190.11	184.95	2.79
	0.3	201.04	201.78	0.37	202.60	201.78	0.40
	0.4	185.12	187.72	1.39	189.01	187.73	0.69
	0.5	163.87	162.69	0.72	165.85	162.69	1.94
	0.6	138.77	137.41	0.99	138.14	137.41	0.53
	0.7	122.40	114.91	6.52	114.28	114.91	0.55
	0.8	108.80	101.82	6.85	102.65	101.82	0.82
	0.9	100.67	93.80	7.33	94.76	93.79	1.03
average		150.27	148.10		149.31	148.10	

As can be seen from the data in Tables 2 and 3, values of activation energy calculated by Ozawa method, KAS method as well as iterative iso-conversional procedure are very approached and the errors of activation energy associated with conventional iso-conversional methods without performing the iterative procedure are less than 10%, indicating the reliability of the computational results. Moreover, compared with average activation energy for each

stage under combustion and gasification conditions, it can be interesting to find that activation energy at dehydration stage is the smallest under both conditions, and the activation energy at pyrolysis stage and char conversion stage under gasification condition are very approached while the activation energy at pyrolysis stage is slightly larger than that under combustion condition.

In order to better analyse large amounts of data in Tables 2 and 3, Fig. 6 is designed to exhibit variation tendency of activation energy with degree of conversion for each stage under combustion or gasification condition. Since the great similarities in activation energy determined by iterative iso-conversional procedure for Ozawa and KAS method, values of activation energy determined by iterative iso-conversional procedure for KAS method are chosen as the data for the Fig. 6.

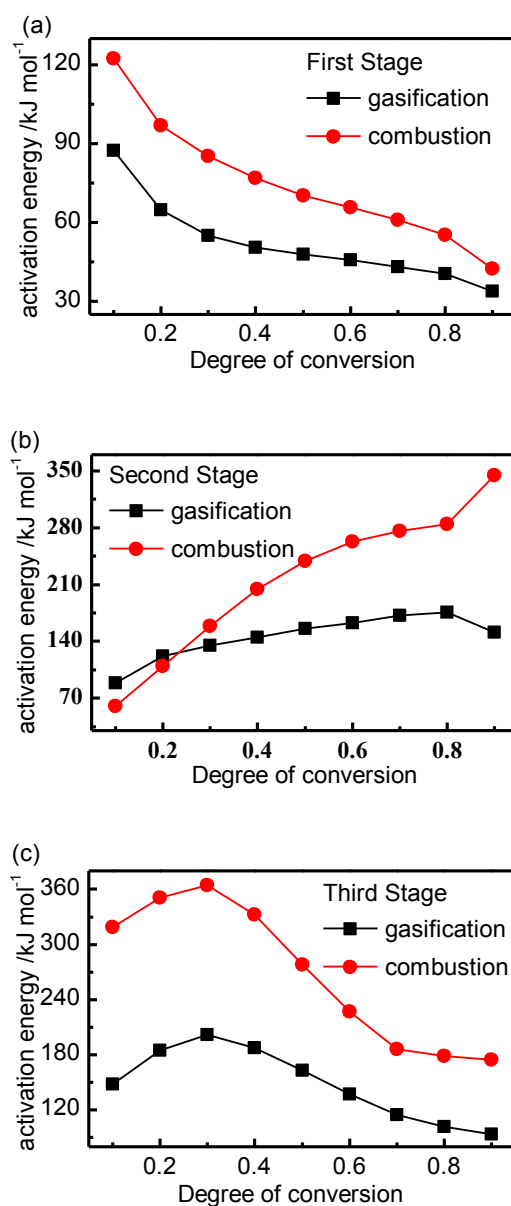


Fig. 6 The variation tendency of activation energy for three stages under combustion and gasification conditions with the increase of degrees of conversion
 (a) First stage-dehydration; (b) Second stage-pyrolysis; (c) Third stage-char conversion

As shown in Fig. 6, it is interesting to find that activation energy under combustion condition is mostly larger than that under gasification condition. The computational result is the same as thermal treatment of charcoal when Boateng et al. [51] investigated thermal conversion of charcoal under combustion and gasification conditions. Moreover, the variation tendency of activation energy with degree of conversion for each stage under combustion or gasification condition shows the same variation. With the increase of degrees of conversion, the values of activation energy under both conditions tend to decrease in the dehydration stage, which is opposite to the variation tendency in the pyrolysis stage. While in the char conversion stage, the values of activation energy tend to increase first and then decrease, reaching a maximum value when degree of conversion is equal to 0.3.

The most probable mechanism functions for each stage under combustion and gasification conditions, as well as the slope and the correlation coefficient of liner regression R^2 , are tabulated in Table 4 according to Eq. 12. As can be seen in Table 4, the slope and the correlation coefficient of liner regression are very near to -1 and 1, respectively, indicating the reliability of the $g(\alpha)$ function [36, 52]. For combustion and gasification conditions, distinguishing $g(\alpha)$ functions were determined for each stage and even at the same stage, $g(\alpha)$ functions differ significantly when different thermal treatments were applied.

Table 4 The most probable mechanism function $g(\alpha)$, slope and the correlation coefficient of liner regression R^2 for each stage under combustion and gasification conditions

stage	combustion process			gasification process		
	$g(\alpha)$	slope	R^2	$g(\alpha)$	slope	R^2
I	$[-\ln(1-\alpha)]^2$	-0.9957	0.9867	$[(1-\alpha)^{-1/3}-1]^2$	-1.0088	0.9999
II	α^2	-1.0164	0.9957	$(1-\alpha)^{-2}-1$	-0.9871	0.9863
III	$1-2\alpha/3-(1-\alpha)^{2/3}$	-1.0185	0.9973	$(1-\alpha)^{-1}-1$	-1.0108	0.9467

Furthermore, due to the approximate solution applied in KAS equation is more accurate than that applied in Ozawa equation [52], the values of pre-exponential factor can be estimated according to determined $g(\alpha)$ functions and the slopes of plots of KAS equation. The computational results of pre-exponential factor, as well as the average values of activation energy for each stage under combustion and gasification conditions, are tabulated in Table 5. Although the average values of activation energy under combustion condition are lightly larger than that under gasification condition, yet the values of pre-exponential factor for each stage under combustion condition are several orders of magnitude than that under gasification condition. Hence, considering the combination effect of activation energy and pre-exponential factor in Eq. 1, it is certain that the reaction rate of thermal decomposition of *P. vittata* under combustion condition is still great larger than that under gasification condition, which is consistent with the actual experimental result.

Table 5 The pre-exponential factor and activation energy for each stage under combustion and gasification conditions

stage	Combustion process		Gasification process	
	A /s ⁻¹	E /kJ mol ⁻¹	A /s ⁻¹	E /kJ mol ⁻¹
I	2.53×10 ⁹	75.14	6.27×10 ²	52.07
II	3.32×10 ¹⁸	215.33	1.19×10 ⁷	145.22
III	1.17×10 ¹⁵	267.82	6.37×10 ³	148.10

4. Conclusion

Thermal behaviour of *P. vittata* was successfully investigated by TG and kinetic triplet was firstly determined according to conventional iso-conversional methods, subsequently modified by an iterative procedure for activation energy. Our results show that:

- 1) Both combustion and gasification can easily realise complete conversion of *P. vittata*.
- 2) Although activation energy for each stage under combustion condition is mostly larger than that under gasification, the reaction rate of thermal decomposition of *P. vittata* under combustion condition is still great larger than that under gasification condition.

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