

**2006 ERSD Annual Report**  
*DOE-BER Environmental Remediation Sciences Project # 95011*

**Technetium Chemistry in High-Level Waste**

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**Research Objectives:**

Tc contamination is found within the DOE complex at those sites whose mission involved extraction of plutonium from irradiated uranium fuel or isotopic enrichment of uranium. At the Hanford Site, chemical separations and extraction processes generated large amounts of high level and transuranic wastes that are currently stored in underground tanks. The waste from these extraction processes is currently stored in underground High Level Waste (HLW) tanks. However, the chemistry of the HLW in any given tank is greatly complicated by repeated efforts to reduce volume and recover isotopes. These processes ultimately resulted in mixing of waste streams from different processes. As a result, the chemistry and the fate of Tc in HLW tanks are not well understood. This lack of understanding has been made evident in the failed efforts to leach Tc from sludge and to remove Tc from supernatants prior to immobilization. Although recent interest in Tc chemistry has shifted from pretreatment chemistry to waste residuals, both needs are served by a fundamental understanding of Tc chemistry.

The fate of Tc in the insoluble salts that constitute tank residuals is of paramount importance due to the long half-life and environmental mobility of Tc. Knowledge gaps include determination of the Tc oxidation state, the stability of reduced Tc solution species, and interactions with possible organic complexants in HLW. The objective of this renewal proposal is to continue to pursue fundamental understanding of Tc solution chemistry that provides the basis to make knowledgeable decisions and predictions of Tc behavior during retrieval operations and in tank residuals.

We propose to complete the following task with carryover funding:

- Investigation of the stability of the Tc-oxalate complex in support of HLW sludge retrieval efforts.

The results of the proposed work will provide the basis for understanding Tc chemistry in HLW that is necessary for waste retrieval processing, accessing the long-term risk of HLW tank residuals, and the design of pretreatment strategies, if necessary.

**Research Progress and Implications**

Significant accomplishments were achieved during the five years of funding. Several of these results have been published or submitted for journal publication. Much of the success we have attained in this research project can be attributed to ability to maintain rigorous reducing conditions over extended periods of time. In our solubility studies of Tc(IV) aqueous solution chemistry, a significant proportion, typically more than 85%, of

Tc in solution is present as Tc(IV). In contrast, previous studies were conducted in the presence of up to 100X excess of Tc(VII). In addition, we have emphasized spectroscopic characterization of solution and solid phase species to validate predictions based on thermodynamic data derived from the solubility studies. The predominance of Tc(IV) solution species has been essential for spectroscopic investigations (UV-Vis and XAS), which greatly facilitate interpretation of our solubility studies.

- We are in the process of conducting a detail experimental and theoretical spectrophotometric study of the  $\text{TcO}(\text{OH})_3^-$  and  $\text{TcO}(\text{OH})_4^{2-}$  species under basic conditions. We first observed these species in our earlier solubility studies of  $\text{TcO}_2 \cdot n\text{H}_2\text{O}(\text{am})$  to 3 M NaOH, where weak amphoteric behavior of Tc(IV) to pH 12 has been observed in the absence of complexing ligands. Analog studies in the presence of variable concentrations of dissolved  $\text{CO}_2$  to understand the formation of Tc(IV)-carbonate complexes at nearly neutral pH conditions had been planned but were not initiated due to termination of the HLW program element.
- Solvent extraction method was used to determine the stability constants of Tc(IV) with oxalate anions in NaCl solutions ranging in concentration from 0.5 M to 2.0 M. The distribution ratio of Tc(IV) between the organic and aqueous phases was found to decrease as the concentration of oxalic acid increased. At the oxalic acid concentrations used in these experiments, the complexes  $\text{TcO}(\text{Ox})$  and  $\text{TcO}(\text{Ox})_2^{2-}$  were found to be the dominant aqueous species. Based on these data, the thermodynamic stability constants of Tc(IV) with oxalate complexes were calculated by the Specific Ion Interaction Theory (SIT). The thermodynamic stability constants of Tc(IV) with oxalate complexes were calculated by the SIT model to be  $\log \beta^{\circ}_{101} = 7.22 \pm 0.17$  for the formation of  $\text{TcOOx}$  ( $\text{TcO}^{2+} + \text{Ox}^{2-} = \text{TcO}(\text{Ox})$ ),  $\log \beta^{\circ}_{102} = 11.08 \pm 0.19$  for the formation of  $\text{TcO}(\text{Ox})_2^{2-}$  ( $\text{TcO}^{2+} + 2\text{Ox}^{2-} = \text{TcO}(\text{Ox})_2^{2-}$ ), respectively. The data are comparable but a little weaker than what has been observed for  $\text{VO}^{2+}$  and  $\text{UO}_2^{2+}$ . No data are available for the stability constants of Tc(IV)-oxalate complexes in the literature. The results indicate that the oxalate anion forms strong complexes with technetium (IV) ion as a function of ionic strength in NaCl solutions. These results will help us to assess the efficiency of tank waste retrieval methods for technetium as well as the long-term mobility in the environment. In order to understand the chemical behavior of technetium with oxalic acid, an integrated approach of solvent extraction, spectroscopic, solubility measurements of  $\text{TcO}_2(\text{am})$  and thermodynamic models in multi-electrolyte systems has been initiated.

### **Planned Activities**

We plan to publish the results of our continuing solubility studies of  $\text{TcO}_2 \cdot n\text{H}_2\text{O}(\text{am})$  as a function of pH in the presence of oxalate under mildly acidic conditions and continued investigations of the structure of Tc precipitates under alkaline conditions. The solubility studies of  $\text{TcO}_2 \cdot n\text{H}_2\text{O}(\text{am})$  in the presence of oxalate are near completion.

### **Information Access**

Three journal articles and one technical report are in various stages of publication. In addition several presentations were made. These are listed below.

### *Publications*

- “Stability constants of Tc(IV) oxalate complexes as a function of ionic strength,” by Y. Xia, N.J. Hess, and A.R. Felmy, *Radiochimica Acta* **94 (3)**: 137-141(2005).
- “A thermodynamic model for the solubility of  $\text{TcO}_2 \cdot x\text{H}_2\text{O}(\text{am})$  in the aqueous Tc(IV)- $\text{Na}^+$ - $\text{Cl}^-$ - $\text{H}^+$ - $\text{OH}^-$ - $\text{H}_2\text{O}$  system,” by N.J. Hess, Y. Xia, D. Rai, and S.D. Conradson. *Journal of Solution Chemistry* 33:199-227, 2004.
- “Solubility of  $\text{TcO}_2 \cdot x\text{H}_2\text{O}(\text{am})$  in the presence of gluconate in aqueous solution,” by N.J. Hess, Y. Xia, and A.R. Felmy, to be published in *American Chemical Society Proceedings Volume entitled “Nuclear Waste Remediation: Accomplishments of the Environmental Management Science Program,” in press.*
- “Bond valence sums for Tc-O system from EXAFS data,” by D.W. Wester and N.J. Hess, *Inorganic Chimica Acta*, **358(4)**:865-874 (2005).
- “Hanford Tanks 241-AY-102 and 241-BX-101: Sludge Composition and Contaminant Release Data,” by Krupka et al. PNNL Technical Report, 2004.

### *Presentations*

- “Solubility of  $\text{TcO}_2 \cdot x\text{H}_2\text{O}(\text{am})$  in the presence of gluconate in aqueous solution.” Invited presentation at 2003 ACS Annual Meeting.
- “Technetium in HLW: An example of environmental science at PNNL.” Invited presentation in Merida, Mexico on May 26<sup>th</sup> 2003.
- “Technetium Chemistry in HLW: Role of Organic Complexants.” 2003 EMSP HLW Annual Workshop.
- “Technetium Chemistry in HLW: Role of Organic Complexants.” FY02 EMSP High Level Waste Workshop.