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**CREVICE REPASSIVATION POTENTIAL OF ALLOY 22 IN HIGH-NITRATE
DUST DELIQUESCENT TYPE ENVIRONMENTS**

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ABSTRACT

The nitrate ion (NO_3^-) is an inhibitor for crevice corrosion of Alloy 22 (N06022) in chloride (Cl^-) aqueous solutions. Naturally formed electrolytes may contain both chloride and nitrate ions. The higher the ratio $R = [\text{NO}_3^-]/[\text{Cl}^-]$ in the solution the stronger the inhibition of crevice corrosion. Atmospheric desert dust contains both chloride and nitrate salts, generally based on sodium (Na^+) and potassium (K^+). Some of these salts may deliquescent at relatively low humidity at temperatures on the order of 150°C and higher. The resulting deliquescent brines are highly concentrated and especially rich in nitrate. Electrochemical tests have been performed to explore the anodic behavior of Alloy 22 in high chloride high nitrate electrolytes at temperatures as high as 150°C at ambient atmospheres. Naturally formed brines at temperatures higher than 120°C do not induce crevice corrosion in Alloy 22 because they contain high levels of nitrate. The inhibitive effect of nitrate on crevice corrosion is still active for temperatures higher than 100°C .

Keywords: N06022, Sodium Chloride, Potassium Nitrate, High Temperature, Deliquescence

INTRODUCTION

Alloy 22 (N06022) is a nickel based alloy designed to be resistant to all forms of corrosion. Alloy 22 contains 56% nickel (Ni), chromium (Cr), molybdenum (Mo), tungsten (W) and iron (Fe). The range in composition is given in ASTM B 575. ¹ Because of its high level of Cr, Alloy 22 remains passive in most industrial environments and therefore it has an exceptionally low dissolution rate. ²⁻⁶ The combined presence of Cr, Mo and W imparts Alloy 22 with high resistance to

localized corrosion such as pitting corrosion and stress corrosion cracking even in hot high chloride (Cl^-) solutions. ⁷⁻¹² Alloy 22 may suffer localized corrosion such as crevice corrosion when it is anodically polarized in chloride containing solutions. ^{8-10,13-15} The presence of nitrate (NO_3^-) and other oxyanions in the solution minimizes or eliminates the susceptibility of Alloy 22 to crevice corrosion. ^{8-10,16-23} The value of the ratio R ($R = [\text{NO}_3^-]/([\text{Cl}^-])$) has a strong effect of the susceptibility of Alloy 22 to crevice corrosion. ¹⁶⁻²³ The higher the nitrate to chloride ratio the stronger the inhibition by nitrate. Most of the work on the susceptibility of Alloy 22 to crevice corrosion was at temperatures of 100°C and below. One study contained the behavior of Alloy 22 at temperatures up to 160°C in highly concentrated calcium based chloride and nitrate brines. ²⁴ It was reported that the inhibiting effect of nitrate for crevice corrosion is also valid up to temperatures of 160°C . ²⁴

Corrosion at the repository site may occur only in the presence of water. There are two possible sources of water that may contact the container at the repository site, seepage from the drift walls; or, deliquescence of salt and dust that may have accumulated during the initial dry periods. Deliquescence of salt mixtures will be the primary source of aqueous solution that may contact and react with waste package surfaces at temperatures above that for boiling of water (during the repository initial thermal period).

Highly soluble salts such as potassium nitrate can produce concentrated solutions that will remain liquid to temperatures higher than 150°C at ambient pressure. ²⁵ Chloride and nitrate salts based on the metals potassium and sodium are contained in dust that may be present at the Yucca Mountain repository site. ²⁶

The majority of the aqueous solutions that are predicted to form on the waste packages due to dust deliquescence will consist of chloride and nitrate salts of sodium and potassium, and, in some cases, calcium.²⁵⁻²⁶ Recent work indicates that deliquescence of these salt mixtures can occur at temperatures in excess of 180°C (Figure 1), even at a low relative humidity of 10-20%.^{25,26}

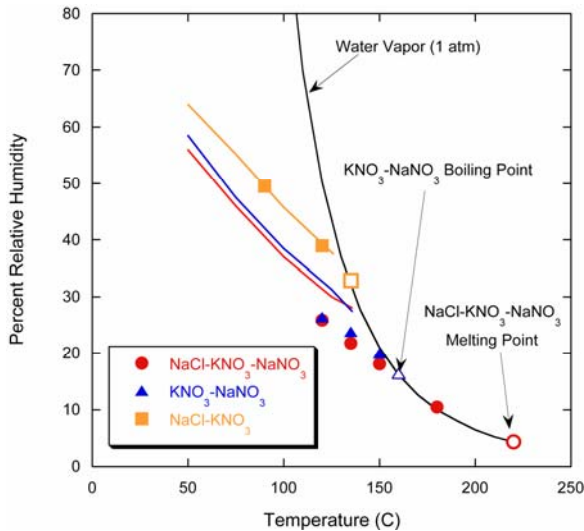


Figure 1. Na-K brines formed by deliquescence²⁵⁻²⁶

The purpose of the current work was to examine the anodic behavior of Alloy 22 in highly concentrated brines at temperatures 100°C and higher. The studies were performed in solutions based mainly on sodium and potassium because these cations are present in the dust. Tests were performed under fully immersed conditions, with unlimited access of brine to the corroding specimen. The volume of concentrated solutions that may form by deliquescence of dust will be small.²⁶

EXPERIMENTAL TECHNIQUE

The specimens of Alloy 22 (N06022) were prism crevice assemblies (PCA) (Figure 2) in the as-welded (ASW) condition. The specimens were machined from welded 1.25-inch thick plates (~32 mm). Table 1 shows the chemical composition of the heats for the base plate and the welding wire. The plates were welded using gas tungsten arc welding (GTAW) from both sides of the plate using the double V groove technique. The exposed surface area for the PCA specimens was 14.06 cm². This surface area did not include the area covered by the crevice formers (CF), which was approximately 1.5 cm². The PCA had a mounting mechanism for the connecting rod explained in ASTM G 5 (Figure 2).²⁷ Crevice formers (CF) were mounted on both sides of the

specimen. Each crevice former consisted of a washer made of a ceramic material containing 12 crevicing spots or teeth with gaps in between the teeth (ASTM G 78).²⁷ Before mounting the CF onto the metallic specimens, they were covered with PTFE tape to ensure a tight crevicing gap between the CF and the specimens. The specimens had a ground surface finish of 600-grit paper.

Table 1. Heat and Composition Of the Tested Specimens

Heat – Manufacturer Of Alloy 22	Composition
Base Plate by Haynes International Heat 2277-0-3183	55.29Ni, 21.23Cr, 13.37Mo, 2.93W, 3.65Fe, 1.7Co, 0.23Mn, 0.14V
Weld Wire by Inco Alloys Heat XX1829BG	59.31Ni, 20.44Cr, 14.16Mo, 3.07W, 2.2Fe, 0.21Mn, 0.15Cu

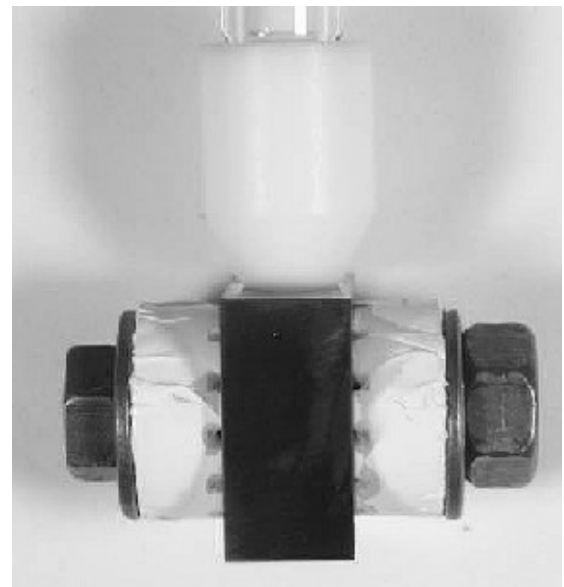


Figure 2. Assembled PCA Specimen

Two different sets of electrochemical tests were performed: (1) Monitoring of the corrosion potential for 24 hours, followed by either a (2) Cyclic Potentiodynamic Polarization (CPP) (ASTM G 61)²⁷ or a (3) Tsujikawa-Hisamatsu Electrochemical (THE) test.¹⁵

Several types of electrolytes were used, from pure chloride to pure nitrate solutions. Table 2 shows a list of tested electrolytes and tested temperatures. Most of the electrolytes were highly concentrated to avoid boiling at ambient pressure. The electrolytes were deaerated with purified nitrogen. The gas stream (N_2) exited the vessel through a condenser to avoid evaporation of the electrolyte. It is not known how the condensation (refluxing) could have affected the pH of the electrolyte. The volume of the solution used for each test was 900 mL. Potentials were monitored using saturated silver chloride electrodes [SSC] through a Luggin capillary. The capillary was filled with 5 M $CaCl_2$ solution because most of the tested solutions are not liquid at ambient temperatures. The reference electrode was kept near room temperature using a jacketed electrode holder through which cooled water was recirculated. The potentials in this paper are reported in the saturated silver chloride scale [SSC]. At ambient temperature, the SSC scale is 199 mV more positive than the normal hydrogen electrode (NHE).

EXPERIMENTAL RESULTS AND DISCUSSION

Evolution of the Corrosion Potential

Figure 3 shows the evolution of the E_{corr} for four Alloy 22 specimens as a function of the immersion time in deaerated brines at $125^\circ C$. Two specimens were immersed in 1 m $CaCl_2$ + 10 m $Ca(NO_3)_2$ ($R = 10$) and two specimens were immersed in 4 m KCl + 21 m KNO_3 + 21 m $NaNO_3$ ($R = 10.5$). The open circuit potential was taken every 20 seconds. The Ca based brine has a concentration by weight (solute/solution) of 64% and the Na + K brine has a concentration of 81%. Even though the ratio R of the concentration of nitrate over chloride was practically the same, Figure 3 shows a completely different behavior of the corrosion potential for Alloy 22 depending if the solution was based on Na + K or based on Ca. The open circuit potential for Alloy 22 in the Ca based brines increased rapidly in the first two hours of exposure and then more gradually for the rest of the exposure time. After only 24-immersion, the E_{corr} of Alloy 22 in the Ca brines is on the order of +600 mV SSC, a potential that can be considered high (Table 2). The total change in potential in this relatively short time of 24 hours is more than 700 mV. In the Na + K brine, the open circuit potential of Alloy 22 slowly decreased as a function of time and after 18-h immersion it was on the order of -450 mV (see also Table 2). The total change since immersion is less than 50 mV. It is not clear at this moment why Alloy 22 behaves so differently in these two types of electrolytes. It is likely that it is related to the hydrolysis properties of the metal cations in the electrolyte, which may control the pH of the electrolyte in contact with the alloy.

Figure 4 shows comparatively the evolution of the open circuit potential for Alloy 22 in pure chloride solutions of the same concentration (8 molal chloride) at $110^\circ C$. Again, as

shown at $125^\circ C$ for the chloride plus nitrate brines (Figure 3), the short-term corrosion potential evolution is different depending if the brine is Ca or Na + K based. In the Ca brine the corrosion potential increased as the time increased and it changed more than 100 mV in 24 hours. In the Na + K brines, the open circuit potential tended to decrease as the immersion time increased. In one of the specimens (KE0691) the decrease in potential was only 10 mV.

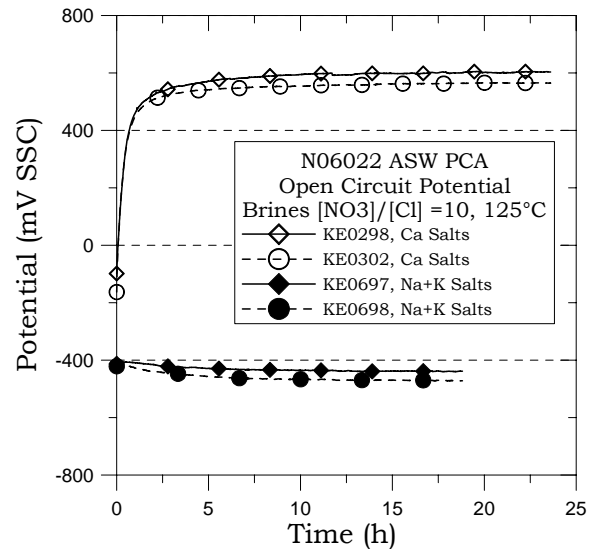


Figure 3. Corrosion potential vs. time in Ca and Na + K brines at $125^\circ C$

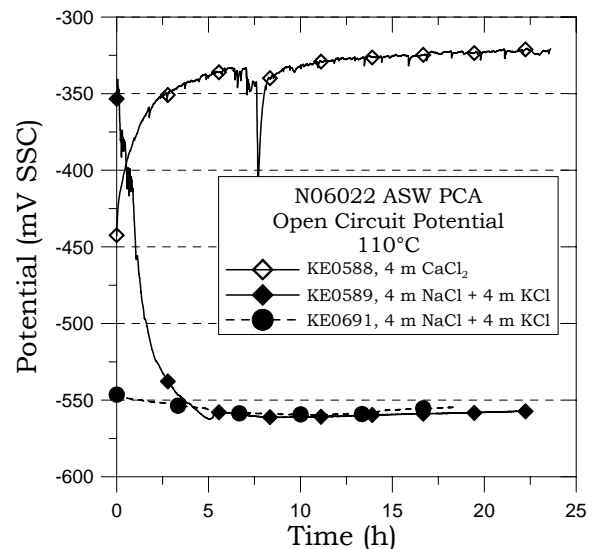


Figure 4. Corrosion potential vs. time in Ca and Na + K brines at $110^\circ C$

Figure 5 shows corrosion potential at the end of the short-time exposure of 24 hours for all the specimens tested at 125°C (Table 2). Figure 5 shows that for even up to $R = 100$, the corrosion potential of Alloy 22 in the Na + K brines was not greatly affected by the amount of nitrate in solution. However, in the Ca based brine, as the ratio R increased the corrosion potential increased. The short-term E_{corr} of Alloy 22 in Ca based brines at 110°C was more than 800 mV higher than the E_{corr} in the Na + K brines.

The values of corrosion potential in Table 2 and in Figures 3-5 are for comparative purposes only and they do not intend to represent the long-term behavior of Alloy 22 in these brines.

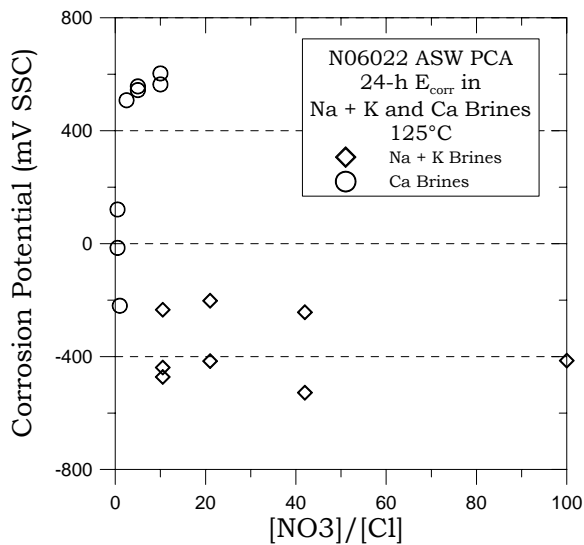


Figure 5. E_{corr} of Alloy 22 in Ca and Na + K brines at 110°C

Effect of Solution Composition on the Cyclic Potentiodynamic Behavior at 110°C

Figure 6 shows the cyclic potentiodynamic polarization (CPP) curves for Alloy 22 in three different electrolyte solutions at 110°C (Table 2). The base concentration of the solution is 8 m chloride (4 m NaCl + 4 m KCl). Figure 6 shows that the breakdown potential increased as the amount of nitrate in the electrolyte solution increased. In the pure Cl solution for a ratio $R = 0$ the breakdown potential (E_{20}) for specimen KE0691 was 49 mV SSC (Table 2). For specimen KE0687 for a solution with a ratio $R = 0.1$, the E_{20} increased to 361 mV SSC and for specimen KE0287 for $R = 0.5$ the E_{20} was 621 mV SSC. The repassivation potential E_{R1} also increased when the amount of nitrate in the electrolyte increased (Table 2). For specimen KE0691 ($R = 0$) E_{R1} was -184 mV SSC, for specimen KE0687 ($R = 0.1$) E_{R1} was -46 mV SSC and for specimen KE0287 ($R = 0.5$) E_{R1} was 327 mV SSC. Figure 6

also shows that both for $R = 0$ and $R = 0.1$ the reverse polarization curve showed a hysteresis, often an indication of localized corrosion (crevice corrosion). When the ratio R was 0.5, the reverse polarization did not show hysteresis, often an indication of absence of localized corrosion. Figure 7 shows that specimen KE0691 after testing in pure 8 molal chloride solution at 110°C, exhibited crevice corrosion. On the other hand, specimen KE0580 tested in a solution containing 8 molal chloride and 4 molal nitrate ($R = 0.5$), was free from crevice corrosion (Figure 8), even though specimen KE0580 was anodically polarized to much higher potentials and for a longer time than specimen KE0691.

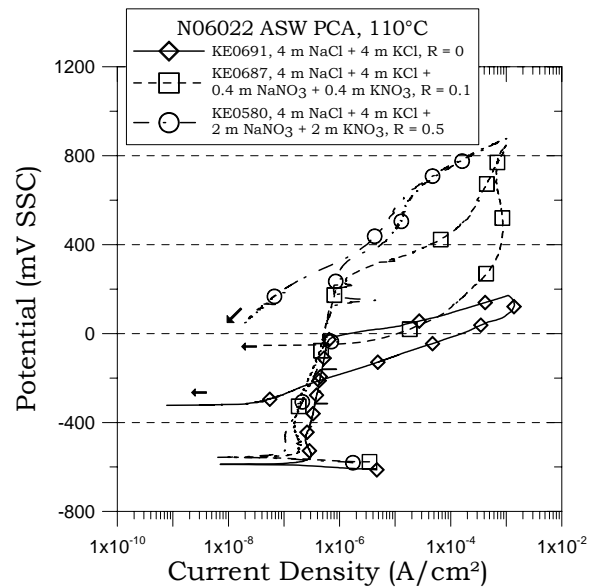


Figure 6. Cyclic Potentiodynamic Polarization of Alloy 22 in Cl^- and $Cl^- + NO_3^-$ brines at 110°C



Figure 7. KE0691 after CPP test Crevice corrosion is observed

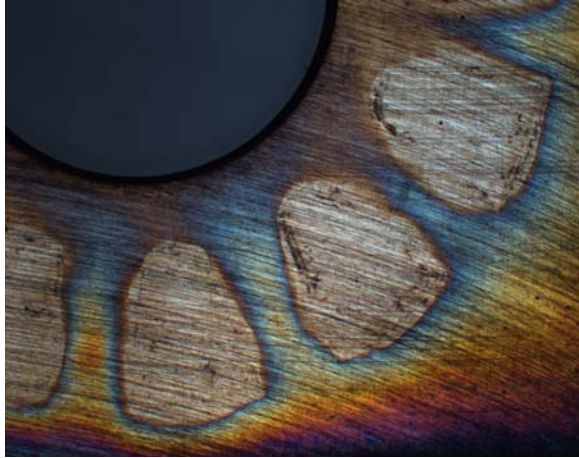


Figure 8. KE0580 after CPP test
No crevice corrosion

Results from Figures 6-8 clearly shows the inhibiting effect of nitrate on the crevice corrosion susceptibility of Alloy 22, even in highly concentrated brines at 110°C.

Cyclic Polarization at Higher Temperatures

Figure 9 shows the anodic behavior of Alloy 22 at 125°C in highly concentrated nitrate solutions. Specimen KE0697 was tested in a brine containing nitrate and chloride to a ratio R of 10.5 and Specimen KE0281 was tested in a pure nitrate solution ($R = \infty$). The behavior of Alloy 22 was practically the same in both solutions. The breakdown potentials E20 and repassivation potentials ER1 were high (775 and 625 mV SSC for KE0697 and 783 and 686 mV for KE0281) (Table 2). The high values of breakdown potentials suggest an increase in current due to the decomposition of water. The reverse scan occurred without hysteresis or with a negative hysteresis suggesting the absence of localized corrosion. Figure 10 shows that both specimens were free of localized corrosion after the tests.

Figure 11 shows the anodic behavior of Alloy 22 at 140°C and 150°C. The behavior of Alloy 22 was practically the same in both solutions and temperatures. The breakdown potentials E20 were 785 mV SSC for KE0700 and 806 mV SSC for KE0284. Both CPP curves exhibited negative hysteresis in the reverse scan. None of the specimens suffered crevice corrosion in spite of the high applied potential (Figure 12).

Repassivation Potentials

Figure 13 shows the repassivation potentials for Alloy 22 at 110°C in pure chloride and low nitrate concentration

electrolytes ($R \leq 1$). There are two type of repassivation potentials represented, one is ER1 from the CPP tests (circles) and the other CRP from the THE tests (squares) (Table 2). The line joins the average values of the ER1 potentials. Figure 13 shows that both type of tests (CPP and THE) yield comparable values of repassivation potentials. The repassivation potential was approximately -200 mV SSC in the pure chloride solution ($R = 0$) and climbed rapidly as the nitrate concentration was increased. At a ratio $R = 0.5$ the repassivation potential was approximately 400 mV SSC, that is, 600 mV higher than in pure chloride solutions. The repassivation potential was higher than 450 mV SSC for a ratio $R = 1$. Figure 13 shows the strong effect of nitrate as an inhibitor of crevice corrosion in Alloy 22.

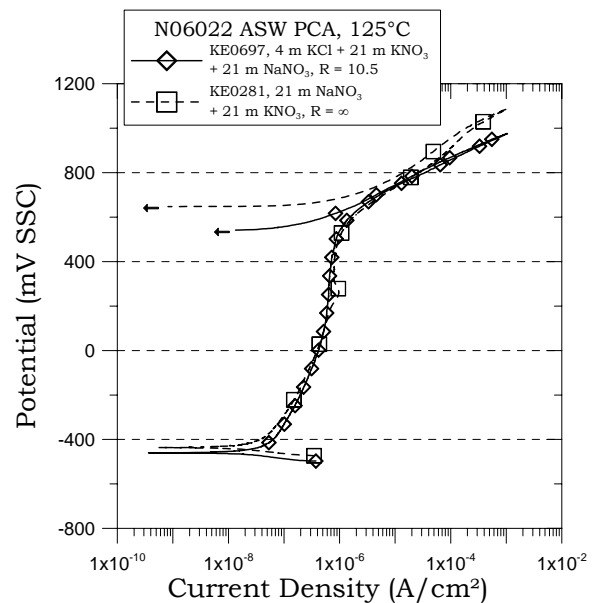
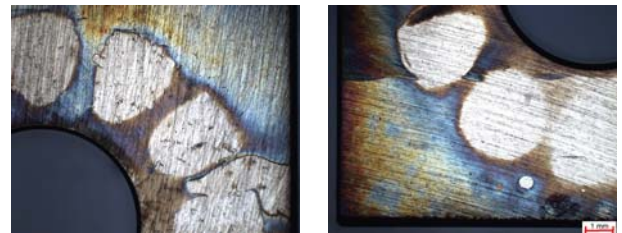


Figure 9. CPP of Alloy 22 in brines at 125°C



KE0697, R = 10.5

KE0281, R = ∞

Figure 10. No localized corrosion after CPP in high nitrate brines at 125°C

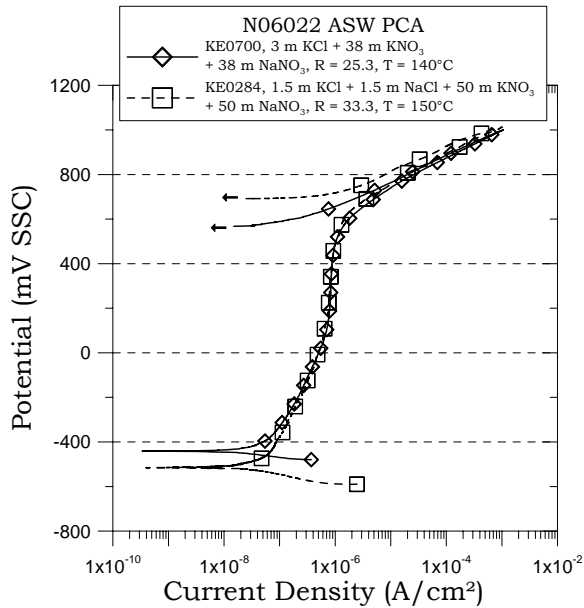


Figure 11. CPP of Alloy 22 at 140°C and 150°C



KE0700, R = 25.3

KE0284, R = 33.3

Figure 12. No localized corrosion after CPP in high nitrate brines at 140°C and 150°C

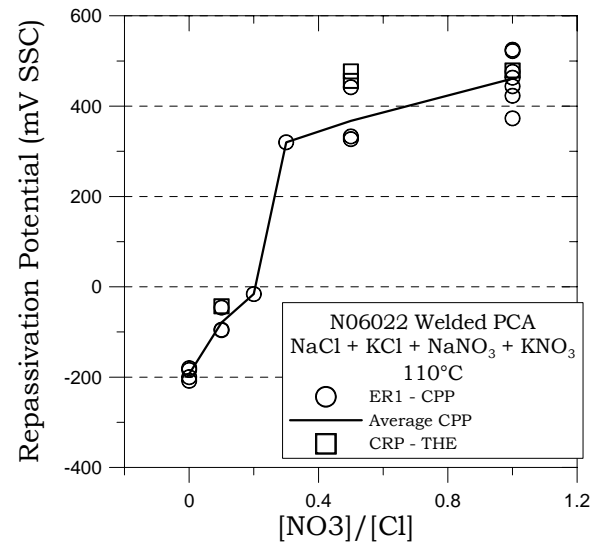


Figure 13. Repassivation potentials, 110°C, $R \leq 1$

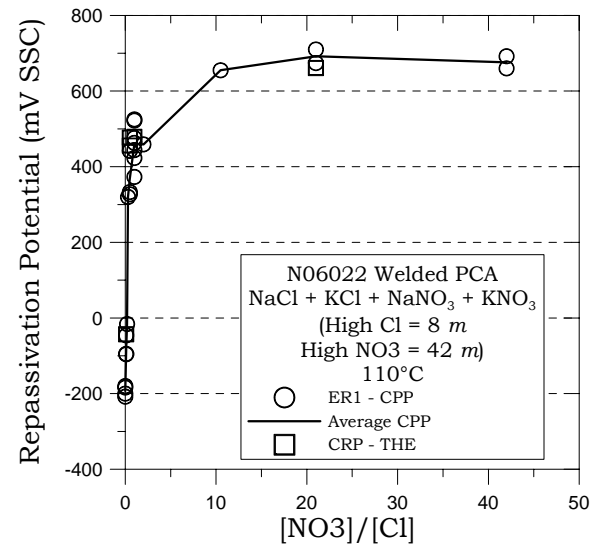


Figure 14. Repassivation potentials at 110°C

Figure 14 shows all the repassivation potentials at 110°C as a function of the ratio R. It is apparent that the repassivation potentials were higher than 600 mV for R values between 10.5 and 42. None of the specimens tested in these high nitrate solutions suffered crevice corrosion (Table 2).

Figure 15 shows the repassivation potentials obtained at 125°C. For R values of 2.5 and higher, the tested specimens were free from localized corrosion. For specimens KE0294 and KE0299 tested in a Ca based brine at an R value of 0.5, pitting corrosion was observed. This may have been the result of the high potential applied to the specimens. In all the cases, the repassivation potentials were high, on the order of 600 mV or higher. Figure 15 shows only repassivation potentials for Ca based brines at R less than 10 because it is not possible to prepare a stable brine at 125°C with an R value less than 10 using only Na and K salts.

Figure 16 shows the repassivation potentials for Alloy 22 at 140°C and 150°C. Solutions that are stable at these temperatures would have values of R higher than 20. It is not possible to have liquid brines at 140°C and 150°C with lower nitrate concentrations. For all the tests the repassivation potentials were higher than 600 mV and none of the specimens suffered localized corrosion (Table 2).

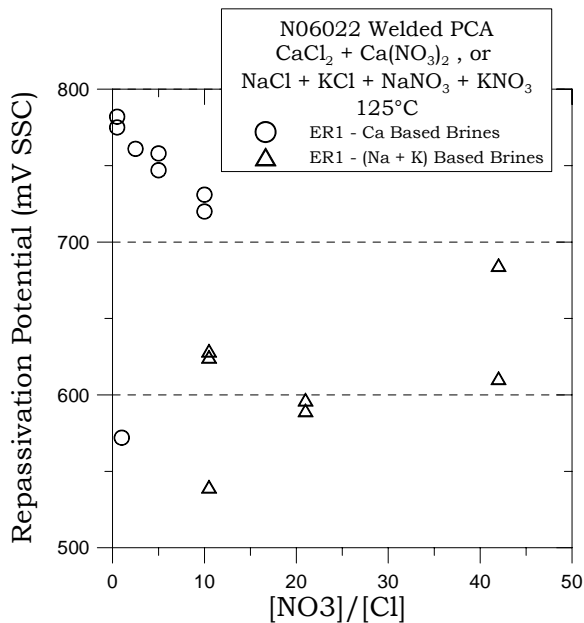


Figure 15. Repassivation potentials at 125°C

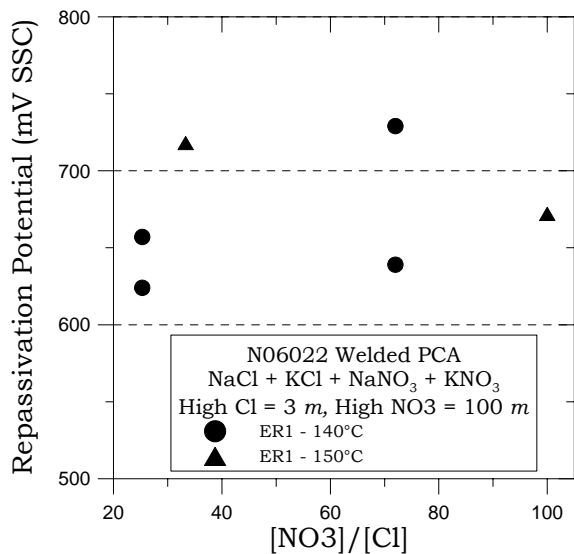


Figure 16. Repassivation potentials at 140°C and 150°C

Concluding Remarks

The dust in the desert near Yucca Mountain may contain salts such as NaCl, KCl, NaNO₃ and KNO₃. Approximately 10 percent of the dust is soluble salts.²⁶ These salts can combine

to form multi-salt assemblages which can deliquesce at temperatures well above the boiling point of water (Figure 1).²⁵ For example a ternary salt assemblage of NaCl, KNO₃ and NaNO₃ would form a liquid solution at ambient pressure at 150°C and 20% relative humidity.²⁵⁻²⁶ These solutions would contain a high concentration in nitrate, with a molar ratio R of nitrate over chloride higher than 1. The higher the temperature the larger the amount of nitrate that needs to be present to cause deliquescence of the salt assemblage. In other words it is not possible to prepare an electrolyte solution that would be rich in chloride at temperatures higher than 120°C.²⁸

Current results confirmed previous findings in Ca based brines regarding the inhibitive effect of nitrate for localized corrosion of Alloy 22 at temperatures higher than 100°C.²⁴ Nitrate will inhibit localized corrosion when it is present in ratios R = 0.5 to 2 and higher. The minimum ratio for total inhibition may be dependent on the temperature and the base concentration of chloride. Brines with values of R lower than 2 can be liquid only at temperatures below approximately 120°C. Current results support the theory that any brine that may form via deliquescence at temperatures higher than approximately 120°C will not induce crevice corrosion in Alloy 22. The current results were obtained by forcing the alloy to corrode by artificial polarization and using a large volume of brine (fully immersed specimens). In the natural emplacement site the amount of brine will be limited and polarization may only occur naturally, for example, via the reduction of oxygen in the air.

CONCLUSIONS

- (1) Nitrate is an inhibitor for localized corrosion at ratios R on the order of 0.5 to 2 and higher
- (2) The inhibitive effect of nitrate is valid at all the tested conditions from 90°C to 150°C.
- (3) Similar values of repassivation potentials were obtained using the cyclic polarization and the THE methods.
- (4) Current results in Na and K brines confirm the same high temperature anodic behavior of Alloy 22 previously reported in Ca brines.

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Nuclear Waste Technical Review Board Workshop on Localized Corrosion, 25-26 September 2006, Las Vegas, Nevada (<http://www.nwtrb.gov/meetings/overheads.html>).

Table 2. Tested Conditions and Characteristic Potentials (in mV, SSC) from the Cyclic Polarization Curves (CPP) and the Tsujikawa-Hisamatsu Electrochemical (THE) tests

Specimen	Solution	Temp. (°C)	$\frac{[NO_3^-]}{[Cl^-]}$	E _{corr}	E20 CPP	E200 CPP	ER10 CPP	ER1 CPP	ERCO CPP	CRP THE	Results
KE0592	4 m KCl + 4 m NaCl + 0.4 m KNO ₃ + 0.4 m NaNO ₃	90	0.1	-123	727	817	684	571	570		CC
KE0688	4 m KCl + 4 m NaCl + 0.4 m KNO ₃ + 0.4 m NaNO ₃	90	0.1	-501	684	805	33	-63	-82		CC
KE0579	4 m KCl + 4 m NaCl + 0.4 m KNO ₃ + 0.4 m NaNO ₃	90	0.1	-482	676	816	83	-53	-80		CC
KE0590	4 m KCl + 4 m NaCl + 0.8 m KNO ₃ + 0.8 m NaNO ₃	90	0.2	-493	695	819	647	405			CC
KE0693	4 m KCl + 4 m NaCl + 0.8 m KNO ₃ + 0.8 m NaNO ₃	90	0.2	-498	686	821	630	384			CC
KE0279	4 m KCl + 4 m NaCl + 0.8 m KNO ₃ + 0.8 m NaNO ₃	90	0.2	-480						N/A	CC
KE0288	4 m KCl + 4 m NaCl + 2 m KNO ₃ + 2 m NaNO ₃	90	0.5	-454						555	CC
KE0689	4 m KCl + 4 m NaCl	110	0	-557	85	141	-123	-200	-223		CC
KE0574	4 m KCl + 4 m NaCl (+ 0.0001 m HCl)	110	0	-413	48	102	-135	-208	-222		CC
KE0691	4 m KCl + 4 m NaCl	110	0	-555	49	115	-101	-184	-210		CC
KE0588	4 m CaCl ₂	110	0	-321	34	67	-123	-180	-243		CC
KE0687	4 m KCl + 4 m NaCl + 0.4 m KNO ₃ + 0.4 m NaNO ₃	110	0.1	-522	361	516	-3	-46	-50		CC
KE0692	4 m KCl + 4 m NaCl + 0.4 m KNO ₃ + 0.4 m NaNO ₃	110	0.1	-396	553	778	-26	-95	-103		CC
KE0573	4 m KCl + 4 m NaCl + 0.4 m KNO ₃ + 0.4 m NaNO ₃ (+ 0.0001 m HCl)	110	0.1	-375	511	738	-39	-96	-106		CC
KE0276	4 m KCl + 4 m NaCl + 0.4 m KNO ₃ + 0.4 m NaNO ₃	110	0.1	-521						-43	CC
KE0589	4 m KCl + 4 m NaCl + 0.8 m KNO ₃ + 0.8 m NaNO ₃	110	0.2	-534	469	746	375	-16	-44		CC
KE0591	4 m KCl + 4 m NaCl + 1.2 m KNO ₃ + 1.2 m NaNO ₃	110	0.3	-562	589	769	481	320	258		No CC
KE0278	4 m KCl + 4 m NaCl + 1.2 m KNO ₃ + 1.2 m NaNO ₃	110	0.3	-500							

Table 2 - Continued - 1

KE0572	4 m KCl + 4 m NaCl + 2 m KNO ₃ + 2 m NaNO ₃ (+ 0.0001 m HCl)	110	0.5	-304	668	780	601	442	446		CC-II
KE0580	4 m KCl + 4 m NaCl + 2 m KNO ₃ + 2 m NaNO ₃	110	0.5	-522	621	794	570	327			No CC
KE0277	4 m KCl + 4 m NaCl + 2 m KNO ₃ + 2 m NaNO ₃	110	0.5	-482						456	CC
KE0289	4 m KCl + 4 m NaCl + 2 m KNO ₃ + 2 m NaNO ₃	110	0.5	-514						477	CC
KE0287	4 m KCl + 4 m NaCl + 2 m KNO ₃ + 2 m NaNO ₃	110	0.5	-491	622		543	333	337		N/A
KE0690	3 m KCl + 3 m NaCl + 3 m KNO ₃ + 3 m NaNO ₃	110	1	-545	576	791	571	373			No CC
KE0571	3 m KCl + 3 m NaCl + 3 m KNO ₃ + 3 m NaNO ₃ (+ 0.0001 m HCl)	110	1	194	651	770	581	444	330		CC-II
KE0575	3 m KCl + 3 m NaCl + 3 m KNO ₃ + 3 m NaNO ₃ (+ 0.0001 m HCl)	110	1	-407	683	759	569	423	413		CC-II
KE0581	3 m KCl + 3 m NaCl + 3 m KNO ₃ + 3 m NaNO ₃	110	1	214	682	779	622	522	496		No CC
KE0285	3 m KCl + 3 m NaCl + 3 m KNO ₃ + 3 m NaNO ₃	110	1	-358	708		633	476	481		N/A
KE0308	3 m KCl + 3 m NaCl + 3 m KNO ₃ + 3 m NaNO ₃	110	1	-485						479	CC-II
KE0291	4 m CaCl ₂ + 4 m KNO ₃ + 4 m NaNO ₃	110	1	-371	700	864	643	463			CC-II
KE0301	4 m CaCl ₂ + 4 m KNO ₃ + 4 m NaNO ₃	110	1	-384	713	867	673	525			No CC
KE0286	3 m KCl + 3 m NaCl + 6 m KNO ₃ + 6 m NaNO ₃	110	2	-431	696		644	459	461		N/A
KE0582	4 m KCl + 21 m KNO ₃ + 21 m NaNO ₃ (+ 0.0001 m HCl)	110	10.5	-106	799	907	764	655	615		No CC
KE0695	2 m KCl + 21 m KNO ₃ + 21 m NaNO ₃	110	21	-233	828	934	779	674	644		No CC
KE0583	2 m KCl + 21 m KNO ₃ + 21 m NaNO ₃ (+ 0.0001 m HCl)	110	21	-296	816	926	760	710	661		No CC
KE0309	2 m KCl + 21 m KNO ₃ + 21 m NaNO ₃	110	21	-382						661	No CC
KE0696	1 m KCl + 21 m KNO ₃ + 21 m NaNO ₃	110	42	-427	845	990	794	692	797		No CC
KE0584	1 m KCl + 21 m KNO ₃ + 21 m NaNO ₃ (+ 0.0001 m HCl)	110	42	-324	829	960	781	660	783		No CC
KE0290	7 m KNO ₃ + 7 m NaNO ₃	110	∞	-373	619		464	393	403		N/A
KE0280	21 m KNO ₃ + 21 m NaNO ₃	110	∞	-361	855	1038	861	772			No CC
KE0294	12 m CaCl ₂ + 6 m Ca(NO ₃) ₂	125	0.5	-15	949	1061	806	782	795		LC - PC

Table 2 - Continued - 2

KE0299	12 m CaCl ₂ + 6 m Ca(NO ₃) ₂	125	0.5	121	946	1054	804	775	781		LC - PC
KE0292	6 m CaCl ₂ + 6 m Ca(NO ₃) ₂	125	1	-220	755	979	687	572			CC-II
KE0293	4 m CaCl ₂ + 10 m Ca(NO ₃) ₂	125	2.5	508	875	991	863	761	728		No CC
KE0297	2 m CaCl ₂ + 10 m Ca(NO ₃) ₂	125	5	557	844	962	840	758	740		No CC
KE0303	2 m CaCl ₂ + 10 m Ca(NO ₃) ₂	125	5	544	837	955	831	747	725		No CC
KE0295	1 m CaCl ₂ + 10 m Ca(NO ₃) ₂	125	10	603	821	922	796	720	689		No CC
KE0302	1 m CaCl ₂ + 10 m Ca(NO ₃) ₂	125	10	564	830	928	804	731	701		No CC
KE0697	4 m KCl + 21 m KNO ₃ + 21 m NaNO ₃	125	10.5	-439	775	894	746	625			No CC
KE0698	4 m KCl + 21 m KNO ₃ + 21 m NaNO ₃	125	10.5	-472	767	888	745	629			No CC
KE0585	4 m KCl + 21 m KNO ₃ + 21 m NaNO ₃ (+ 0.0001 m HCl)	125	10.5	-234	743	857	690	540	538		No CC
KE0586	2 m KCl + 21 m KNO ₃ + 21 m NaNO ₃ (+ 0.0001 m HCl)	125	21	-202	776	893	733	597	629		No CC
KE0587	2 m KCl + 21 m KNO ₃ + 21 m NaNO ₃	125	21	-416	765	885	720	590	715		No CC
KE0686	1 m KCl + 21 m KNO ₃ + 21 m NaNO ₃	125	42	-528	749	911	734	611			No CC
KE0576	1 m KCl + 21 m KNO ₃ + 21 m NaNO ₃ (+ 0.0001 m HCl)	125	42	-243	785	994	762	685	847		No CC
KE0281	21 m KNO ₃ + 21 m NaNO ₃	125	∞	-414	783	982	779	686			No CC
KE0699	3 m KCl + 38 m KNO ₃ + 38 m NaNO ₃	140	25.3	-132	784	903	748	624	651		No CC
KE0700	3 m KCl + 38 m KNO ₃ + 38 m NaNO ₃	140	25.3	-418	785	912	764	657			No CC
KE0577	1 m KCl + 36 m KNO ₃ + 36 m NaNO ₃	140	72	-384	818	997	826	729			No CC
KE0578	1 m KCl + 36 m KNO ₃ + 36 m NaNO ₃ (+ 0.0001 m HCl)	140	72	-352	786	970	767	639			No CC
KE0284	1.5 m KCl + 1.5 m NaCl + 50 m KNO ₃ + 50 m NaNO ₃	150	33.3	-448	806	933	815	718			No CC
KE0595	0.5 m KCl + 0.5 m NaCl + 50 m KNO ₃ + 50 m NaNO ₃	150	100	-335	785	969	789	672			No CC

In the concentration of the electrolyte "m" stands for molal, which is moles of the solute (salt) per kilogram of solvent (water). CC = Crevice Corrosion. CC-II = Dull type of CC. LC = Localized Corrosion, PC = Pitting Corrosion, E_{corr} = is the free corroding potential after ~24 h exposure in the given deaerated electrolyte. E₂₀ and E₂₀₀ are the potentials in the forward scan of a cyclic potentiodynamic polarization (CPP) curve when the current density first reaches 20 μA/cm² and 200 μA/cm². . ER₁₀ and ER₁ are the potentials in the reverse polarization where the current density first reaches 10 μA/cm² and 1 μA/cm². ER_{CO} is the potential where the reverse scan crosses-over the forward scan. CRP is the repassivation potential using the Tsujikawa-Hisamatsu Electrochemical (THE) method.