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P. Sollins, C. Swanston, M. Kleber, T. Filley, M. Kramer, S. Crow, B. Caldwell, K. Lajtha, R. Bowden

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5 **Organic C and N stabilization in a forest soil: evidence from sequential**
6 **density fractionation**

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8 **Phillip Sollins^{a,*}, Christopher Swanston^b, Markus Kleber^c, Timothy Filley^d, Marc Kramer^e,**
9 **Susan Crow^a, Bruce A. Caldwell^a, Kate Lajtha^a, and Richard Bowden^f**

10 ^aOregon State University, Corvallis, OR 97331, USA

11 ^bLawrence Livermore National Laboratory, Livermore, CA 94550, USA

12 ^cLawrence Berkeley National Laboratory, Berkeley, CA 94720, USA

13 ^dPurdue University, West Lafayette, IN 47907, USA

14 ^eUniversity California Santa Cruz, Santa Cruz, CA 95064, USA

15 ^fAllegheny College, Meadville, PA, 16335, USA
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17

18 **Abstract**

19
20 In mineral soil, organic matter (OM) accumulates mainly on and around surfaces of silt- and
21 clay-size particles. When fractionated according to particle density, C and N concentration (per
22 g fraction) and C/N of these soil organo-mineral particles decrease with increasing particle
23 density across soils of widely divergent texture, mineralogy, location, and management. The
24 variation in particle density is explained potentially by two factors: 1) a decrease in the mass
25 ratio of organic to mineral phase of these particles, and 2) variations in density of the mineral
26 phase. The first explanation implies that the thickness of the organic accumulations decreases
27 with increasing particle density. The decrease in C/N can be explained at least partially by
28 especially stable sorption of cationic peptidic compounds (amine, amide, and pyrrole) directly to
29 mineral surfaces, a phenomenon well documented both empirically and theoretically. These
30 peptidic compounds, along with ligand-exchanged carboxylic compounds, could then form a
31 stable inner organic layer onto which less polar organics could sorb more readily than onto the
32 highly charged mineral surfaces (“onion” layering model).
33

34 To explore mechanisms underlying this trend in C concentration and C/N with particle density,
35 we sequentially density fractionated an Oregon andic soil at 1.65, 1.85, 2.00, 2.28, and 2.55 g
36 cm⁻³ and analyzed the six fractions for measures of organic matter and mineral phase properties.
37

38 All measures of OM composition showed either: (1) a monotonic change with density, or (2) a
39 monotonic change across the lightest fractions, then little change over the heaviest fractions.
40 Total C, N, and lignin phenol concentration all decreased monotonically with increasing density,
41 and ¹⁴C mean residence time (MRT) increased with particle density from ca. 150 y to >980 y in
42 the four organo-mineral fractions. In contrast, C/N, ¹³C and ¹⁵N concentration all showed the
43 second pattern. All these data are consistent with a general pattern of an increase in extent of
44 microbial processing with increasing organo-mineral particle density, and also with an “onion”
45 layering model.
46

1 X-ray diffraction before and after separation of magnetic materials showed that the sequential
2 density fractionation isolated pools of differing mineralogy, with layer-silicate clays dominating
3 in two of the intermediate fractions and primary minerals in the heaviest two fractions. There
4 was no indication that these differences in mineralogy controlled the differences in density of the
5 organo-mineral particles in this soil. Thus, our data are consistent with the hypothesis that
6 variation in particle density reflects variation in thickness of the organic accumulations and with
7 an “onion” layering model for organic matter accumulation on mineral surfaces. However, the
8 mineralogy differences among fractions made it difficult to test either the layer-thickness or
9 “onion” layering models with this soil. Although sequential density fractionation isolated pools
10 of distinct mineralogy and organic-matter composition, more work will be needed to understand
11 mechanisms relating the two factors.

12
13 *Keywords:* soil organic matter, particle density, ^{13}C , ^{15}N , ^{14}C , MRT, lignin, hydroxyfatty acids,
14 cutin, suberin, X-ray diffraction, glycoprotein, amino sugar, protein, amino acid, smectite,
15 magnetite, acid/aldehyde ratio, carbon, nitrogen, sulfur, soil C stabilization.

17 18 **1. Introduction**

19
20 Understanding controls on soil organic matter (SOM) storage is critical to quantifying changing
21 rates of C cycling and storage in response to global climate change. Additionally, SOM plays
22 critical roles in governing agricultural and forest soil dynamics and ecosystem productivity. Soil
23 organic matter (OM) dynamics are complex, requiring understanding of biological and chemical
24 interactions between organic compounds and mineral soil constituents. In mineral soil, OM
25 accumulates mainly on and around surfaces of silt- and clay-size particles. When fractionated
26 according to density, C and N concentration (per g fraction) and C/N of these soil organo-
27 mineral particles decrease with increasing particle density across soils of widely divergent
28 texture, mineralogy, location, and management (Fig. 1). Portions of this trend in C/N are easily
29 understood. The lightest fractions ($<1.6\text{-}1.8\text{ g cm}^{-3}$) have high C/N due simply to the high
30 content of plant-derived polysaccharides, lignin, cutin, suberin, and charcoal. The continued
31 decrease in C/N above densities $> 2.0\text{ g cm}^{-3}$ was commented on by Oades (1989) but remains
32 unexplained.

33
34 To understand this trend in C/N, it is helpful to replot the data underlying Fig. 1 to show particle
35 density as a function of C content (Fig. 2). The strong inverse relation suggests that much of the
36 variation in density of the soil particles is due to differences in C concentration. Alternatively,
37 some could also be due to variation in density of the mineral or organic phases. The first
38 explanation, however, implies that the thickness of the organic accumulations decreases with
39 increasing particle density. Thus it may be reasonable to recast the trend in C/N with density as
40 reflecting a decrease in C/N with decreasing thickness of the organic layer accumulated on
41 mineral surfaces. Layering, perhaps, is not the most the best term to describe this phenomenon,
42 as there is ample evidence that the organic matter is distributed discontinuously over the mineral
43 surfaces at scales $\ll 1\text{ }\mu\text{m}$ (Mayer, 1999; Kaiser and Guggenberger, 2003; Zimmerman et al.,
44 2004). However, scanning electron microscopy at scales $> 1\text{ }\mu\text{m}$ shows very few clean mineral
45 surfaces in the A horizons of most mineral soils (e.g., Spycher et al., 1986), thus we use “layer”
46 here for convenience.

1
2 A low C/N in thin organic layers on mineral surfaces can be explained at least partially by
3 especially stable sorption of cationic nitrogenous compounds (glycoproteins, amino sugars, and
4 proteins) directly to mineral surfaces, a phenomenon well documented both empirically and
5 theoretically. Much soil organic N occurs as peptidic compounds (amide, amine, and pyrrole N;
6 Schulten and Schnitzer, 1998, Knicker, 2000; Knicker et al., 2000; DiCosty et al., 2003; Martens
7 and Loeffelmann, 2003; Schmidt-Rohr et al., 2004). A significant portion may occur as
8 heterocyclic N (Mertz et al. 2005; Smernik and Baldock, 2005), which can be biotic (e.g.,
9 nucleic acids), combustion products associated with charcoal, or fossil organics inherited from
10 sedimentary rock particles. The amount of heterocyclic N in soils is still unclear, however.

11
12 There is considerable evidence that peptidic compounds bind especially stably to mineral
13 surfaces (Theng, 1979, Chevallier et al., 2003; Wershaw, 2004). Strong sorption of defined
14 protein on clean clays is well documented, although values vary over nearly five orders of
15 magnitude, related apparently to the nature of the sorbate, sorbent, and experimental conditions
16 (Theng, 1979, Burchill et al., 1981; Wang and Lee, 1993; De Cristoforo and Violante, 2001;
17 Ding and Henrichs, 2002). The highest levels reported for clean clays (e.g., Burchill et al., 1981,
18 p. 269; DeCristoforo and Violante, 2001) and buried ceramics (Craig and Collins, 2002)
19 approach 1 g protein/g clay, implying a soil C concentration greater than that allowed in Soil
20 Taxonomy for a mineral soil horizon (Soil Survey Staff, 1999). In addition, considerable
21 empirical data for marine sediments and riverine particles show strong and stable binding of
22 amino acids to clay minerals (Hedges and Hare, 1987; Wang and Lee, 1993; Mayer et al., 1995;
23 Arnarson and Keil, 2001; Aufdenkampe et al., 2001). These peptidic compounds could then
24 form a stable inner organic layer onto which less polar organics could sorb more readily than
25 onto the highly charged mineral surfaces, as well as a reactive chemical template for further
26 coupling and cross-linking of organic compounds. We refer to this idea as an “onion” layering
27 model, evidence for which has been obtained by Murphy et al. (1990) who showed that sorption
28 of acidic organics on clean clays greatly increased subsequent sorption of hydrophobic
29 compounds.

30
31 Here we explore changes in SOM and the validity of our onion-layering model through a range
32 of densities in an andic forest soil from the Oregon Cascade range. We isolated physical soil
33 fractions using sequential density fractionation (SDF), then analyzed each fraction for measures
34 of both organic matter composition (total C and N; ^{14}C ; stable C and N isotopes; lignin phenol
35 content and oxidation state; hydroxyl fatty acids as a measure of cutin and suberin) and
36 mineralogy (X-ray diffraction and NaF reaction).

37 38 39 **2. Methods**

40 41 *2.1. Site and soil description*

42
43 Soil was collected adjacent to the Detritus Input and Removal Treatments (DIRT) study plot #18
44 (Lajtha et al. in press) at the Andrews Experimental Forest in western Oregon. The soil, an
45 Andic Dystrudept (J. Dixon and J. Noller, Oregon State University, pers. communic.), is
46 developed in an alluvial/colluvial fan that now forms a gently sloping surface about 20 m above

1 Lookout Creek. Rocks upslope (the presumed source of the deposit) are mainly ash-flow tuffs,
2 both welded and non-welded (Swanson and James 1975), andesitic/basaltic in chemistry. The
3 site is well west (upwind) of the active Cascade volcanoes, thus ash fall has not been a
4 significant factor. The climate is Mediterranean with warm, dry summers and cool, wet winters.
5 Mean annual temperature at the nearby headquarters site averaged 8.8°C and annual precipitation
6 was 2200 mm during 1974-2003 (Sulzman et al., 2005).

7
8 The forest overstory is mixed old-growth Douglas-fir (*Pseudotsuga menziesii*), western hemlock
9 (*Tsuga heterophylla*) and western redcedar (*Thuja plicata*). The Douglas-fir originated in a
10 stand-replacement fire ca. 1500. The overstory is relatively sparse (25 dominants ha⁻¹),
11 permitting good growth of understory species including western hemlock, Pacific yew (*Taxus*
12 *brevifolia*), vine maple (*Acer circinatum*), Pacific dogwood (*Cornus nuttallii*), huckleberry
13 (*Vaccinium* spp.), and sword fern (*Polystichum munitum*). Mosses are abundant in some areas
14 but not where we sampled.

15 16 17 2.2. Sample collection and handling

18
19 For sequential density fractionation, a single large A-horizon sample was collected (July 2004)
20 just outside the SW corner of DIRT plot 18 (Lajtha et al., in press) in an area that had not been
21 disturbed during plot installation. The O horizon (including all recognizable coarse wood
22 fragments) and the upper 2 cm of mineral horizon soil were discarded to decrease the amount of
23 organic debris in the sample. Soil was returned immediately to the laboratory, sieved by pushing
24 the soil through a 2-mm screen, mixed thoroughly, and stored in a cooler (2°C) for several days
25 until fractionation. This large sample was then subsampled for laboratory analyses. Given the
26 sample size ($n=1$), any variability reported thus represents effects of subsampling and lab
27 procedures, not spatial variability at the site.

28 29 2.3. Sequential density fractionation (SDF)

30
31 Density fractionation involved centrifugation of soil suspended in sodium polytungstate (SPT),
32 aspiration of the floating material and supernatant, and vacuum filtration (Strickland et al., 1992;
33 Swanston et al., 2005). First, three 15-g samples were set aside to determine moisture content
34 (26%). Then, eight subsamples (30 g fresh weight) were placed in 225-ml polycarbonate
35 centrifuge tubes with conical bottoms and 5-cm diameter mouths. Such tubes allow excellent
36 physical separation of sediment and floating material, while their transparency facilitates
37 aspiration of the supernatant without disturbing the sediment. Four-place tube racks and vacuum
38 manifolds greatly sped operations. Samples were shaken 2-3 hr on a shaker table. The intent
39 was to disperse only the weakly bound aggregates. The soil suspensions were then centrifuged at
40 2400 rpm (970 g) in a swinging-bucket rotor for 20-30 min, after which floating material was
41 aspirated and rinsed with deionized H₂O on a glass fiber filter (Whatman GF/F, 0.7 μm particle
42 retention). This procedure was done twice at each density, and the aspirated materials from both
43 cycles for a given density were combined. The density of the remaining SPT solution was then
44 increased by adding the next density SPT solution to the centrifuge tubes. Density was assumed
45 to be somewhat light for cycle 1, because of dilution by residual lighter medium. We initially
46 selected target densities of 1.65, 1.85, 2.00, 2.25, and 2.50 g cm⁻³ based on previous studies and

1 known densities of typical soil primary and secondary minerals (see below). The densities
2 actually reported (1.65, 1.85, 2.00, 2.28, and 2.55 g cm⁻³) were those measured at the end of
3 cycle 2 by pipetting and weighing 1 mL of supernatant remaining above the sediment after
4 aspiration. Densities for a given fraction were usually within 0.03 g cm⁻³ of the target value
5 (never more than 0.05 g cm⁻³ different). Fractions from the 8 subsamples were composited
6 randomly (same choice of subsamples for all density fractions) into two samples for subsequent
7 laboratory analysis. All laboratory replicates agreed closely; only the means are reported here.
8 To provide more material, a second density fractionation was done in January 2005 after the
9 composite soil sample had been stored field moist at 2°C for about 6 months.

11 2.4. Organic matter analyses

13 Organic C and N concentrations were determined in triplicate on a Leco CNS-2000 elemental
14 analyzer. Sample size was 0.1-0.2 g for the three lightest fractions, and 0.4-0.5 g for the three
15 heaviest fractions. Samples were also analyzed for δ¹³C and δ¹⁵N with a 20/20 ANCA GSL
16 continuous-flow isotope-ratio mass spectrometer (PDZ Europa, Cheshire, UK) at the Rosenstiel
17 School of Marine Science, University of Miami.

19 Density fractions were analyzed for lignin (Hedges and Mann, 1979) and cutin and suberin acids
20 (hydroxyl and fatty acids and diacids: Goñi and Hedges, 1990; Rumpel et al., 2004) via cupric-
21 oxide (CuO) oxidation. Briefly, small amounts of sample (enough to yield approximately 1-3
22 mg C) were weighed into Monel reaction vessels (Prime Focus) with CuO powder and ferrous
23 ammonium sulfate, and purged under Ar gas for at least 4 hours. The vessels were then filled
24 with 2 M NaOH, placed in a rotating carousel, heated to 150°C (4.2°C min⁻¹) and held for 150
25 minutes (total time: 3 hours). After heating, the samples were opened and ethyl vanillin and DL-
26 12 hydroxystearic acid were added as internal recovery standards (IRS). The high molecular-
27 weight, acid-insoluble organics were then precipitated by acidification, and soluble lignin-
28 derived phenols and functionalized fatty acids were extracted into diethyl ether, blown to
29 dryness, re-diluted in 100 µL pyridine, and stored at -20°C until further processing.

31 Lignin phenols were quantified by analysis of the trimethylsilane (TMS) derivatives of eight
32 lignin phenol monomers (vanillin, acetovanillone, syringaldehyde, vanillic acid,
33 acetosyringone, syringic acid, *p*-hydroxycinnamic acid, and ferulic acid) using a 5 point
34 calibration curve for each of the target compounds relative to the IRS ethyl vanillin. The
35 trimethylsilane (TMS) derivatives of nine hydroxy fatty acids and fatty di-acids were analyzed
36 based upon target and proxy standards relative to the IRS of DL-12, hydroxystearic acid in a
37 five-point calibration curve using extracted ion profiles. Total cutin and suberin acids are
38 defined here as the sum of 16-hydroxyhexadecanoic acid, hexadecanoic diacid, 18-
39 hydroxyoctadec-9enoic acid, and semiquantitation of 9,16&10,16 dihydroxyhexadecanoic acid,
40 9-octadecene1,18dioic acid, 7&8 hydroxyhexadecane dioic, 9,10,18 trihydroxyoctadec12enoic,
41 and 9,10,18-trihydroxyoctanoic acid. A Shimadzu QP5050A quadrupole mass spectrometer
42 interfaced to a GC17a gas chromatograph was used to identify and quantify selected mass
43 fragments of the IRS and target compounds. Successful TMS derivatization of samples and
44 GC/MS instrument performance was verified with a methyl 3,4-dimethoxybenzoate recovery
45 standard, which was added prior to derivatization of all samples and blanks.

1 The lignin analyses provide information on the oxidation state of the lignin compounds and the
2 nature of the plant inputs to the soil. For example, information about the oxidation state of lignin
3 can be inferred from the relative abundance of vanillic acid to vanillin (Ac/Al)_v and syringic acid
4 to syringaldehyde, as microbial degradation of lignin increases the ratio of carboxylic acids to
5 aldehydes (Hedges et al., 1988; Filley et al., 2000). Additionally, the nature of plant inputs to the
6 soil can be inferred by comparing the yield of syringyl or cinnamyl compounds normalized to
7 yield of vanillyl compounds (S/V and C/V, respectively) (Hedges and Mann 1979).

8
9 Radiocarbon was measured on the Van de Graaff FN accelerator mass spectrometer (AMS) at
10 the Center for Accelerator Mass Spectrometry at Lawrence Livermore National Laboratory, CA.
11 Samples were prepared for AMS analysis by combustion to CO₂ in the presence of CuO and
12 subsequent reduction of the CO₂ onto iron powder in the presence of H₂ (Vogel et al., 1987).
13 Radiocarbon data are expressed according to Stuiver and Polach (1977) as $\Delta^{14}\text{C}$, the deviation in
14 parts per thousand from the absolute international standard activity (¹⁴C:¹²C ratio of oxalic acid
15 corrected for decay since 1950). The $\Delta^{14}\text{C}$ values were adjusted for mass-dependent
16 fractionation based on measured $\delta^{13}\text{C}$ values for each fraction (Table 1). Mean residence times
17 of density fractions were calculated with a time-dependent steady-state model (Trumbore, 1993;
18 Trumbore et al., 1995, Gaudinski et al., 2000, Torn et al., 2002). The two fundamental
19 assumptions of the model are: 1) inputs equal losses for each pool through time, though $\Delta^{14}\text{C}$
20 varies; and 2) the $\Delta^{14}\text{C}$ of inputs to all pools is equal to that of the atmosphere in the previous
21 year. Thus, the model does not account for multi-year lags before input or transfer between
22 pools, so resulting MRTs should be considered maximum values. Yearly atmospheric values
23 ($\Delta^{14}\text{C}$) used in the model were based on three chronologies, beginning in calendar year 1511
24 (Stuiver et al., 1998; Hua and Barbetti, 2004; Levin and Kromer, 2004).

25 26 *2.5. Mineralogical analyses*

27
28 NaF reaction pH was measured 2 minutes after adding 1.0 M NaF to 0.3 to 0.6 g of fraction
29 (depending on amounts available) to yield a 1:50 soil:solution ratio. Qualitative identification of
30 soil mineralogy in bulk soil and density fractions was based on X-ray diffraction. Samples were
31 hand-ground with a mortar and pestle to a mean particle size <200 μm , then back-loaded into a
32 2.5-cm diameter circular-cavity holder and run on a PANalytical X'Pert Pro instrument using
33 Co-K α radiation at 40 kV and 40 mA. Diffraction patterns were recorded by step scanning from
34 3-100° 2 θ , with the sample spinning at 2 revolutions per second. Phase identification and peak
35 area determinations were done using the X'pert High Score plus software.

36 37 38 **3. Results**

39 *3.1. Physical appearance, dry weight, and C and N*

40 Fractions are referred to hereafter as F1 through F6 as indicated in Table 1. Light microscopy
41 (Fig. 3) showed that fibrous material and charcoal decreased in abundance steadily with
42 increasing density and were nearly absent from fractions F3-F6. Sand- and coarse silt-size
43 mineral fragments were largely restricted to F5 and F6.

44

1 Concentrations of C and N decreased with increasing particle density from F1 to F6 (Table 1).
2 Dry weight, as percent of total soil material, peaked strongly in F4 (Fig. 4). Although F1
3 comprised only a small portion of the total soil weight, it accounted for more of the total soil C
4 than any other fraction (Fig. 4), a result of its high C concentration (Table 1). Conversely, F4
5 had a relatively low C concentration, but represented a large percentage of the soil mass, and
6 thus also accounted for a large proportion of the total soil C. Nitrogen concentration generally
7 dropped with increasing density, except between F1 and F2. F3 and F4 accounted for the bulk of
8 total soil N.
9

10 *3.2. Mineralogy of the bulk soil and density fractions*

11
12 The bulk soil was dominated by three major phases: plagioclase feldspar, quartz, and smectite
13 (Fig. 5). Traces of pyroxene (enstatite) and a zeolite mineral (possibly heulandite) were also
14 discernible. Peaks representing pedogenic oxides were either not present or concealed by others.
15 There was neither the elevated background usually related to the presence of amorphous glass (as
16 would be seen in a young Andisol) nor the typically very broad ‘hump-like’ signals representing
17 short-range order components (as would be expected in a soil with high allophane content).
18 Generally, the XRD trace gave the impression of a volcanic soil with little influence of volcanic
19 glass and at a weathering stage in which smectite is accumulating but pedogenic oxides have not
20 yet grown to the sizes of better refracting, larger crystals.
21

22 Broad, low intensity signals typical for smaller and less crystalline minerals prevailed in F1 and
23 F2, suggesting that particle sizes and reflecting planes were smaller in these than in the other
24 density fractions. Despite their low signal intensity, the presence of smectite and plagioclase
25 feldspar could be confirmed. The quartz 100 reflection was visible as a sharp peak at 3.34 Å.
26 Both patterns clearly indicated a much lower crystallinity of the mineral matrix in F1 and F2 than
27 in the heavier fractions. Smectite particles in F1 and F2 were assumed to be smaller than those in
28 the heavier fractions (indicated by the “line broadening” leading to very broad peaks or mere
29 humps).
30

31 F3 and F4 were dominated by the phyllosilicate clay mineral smectite (at 6° 2θ in the bulk soil
32 and at 8° 2θ in the density fractions), producing a peak that greatly overshadowed those of the
33 quartz and plagioclase signals at 3.34 and 3.21 Å (Fig. 5). As the medium used for the density
34 separation was a sodium salt (Na polytungstate), the smectite in the density fractions had been
35 transformed to the Na form, which was expressed in a contraction from 15.4 to 12.4 Å
36 (comparing bulk soil with the density fractions). F4 also showed the strongest zeolite signal of all
37 density fractions, and exclusively contained two mineral phases refracting at 9.7-9.9 Å and 10.7
38 Å. It is known that the interlayer space of Na-smectite varies with relative humidity (Moore and
39 Hower, 1986) leading to discrete thicknesses of 9.6, 12.4, 15.2 and 18 Å. We thus attribute the
40 9.7 signal observed here to a “drier” (and, therefore, slightly heavier) portion of the Na-smectite,
41 which seems to have been specifically isolated by the 2.0 g cm⁻³ density cutoff. The signal at
42 10.7 Å was tentatively interpreted as representing the early stages of a smectite-illite transition.
43 F5 and F6 were rich in primary minerals but lacked phyllosilicates. F6 was rich in plagioclase
44 feldspar and magnetite, whereas F5 contained abundant quartz but little magnetite.
45

1 NaF pH was moderately high in all the heavier fractions (Table 1) indicating the presence of
2 considerable amorphous Al or Fe. It rose steadily with density up to F3, then remained
3 essentially constant. Magnetic separation showed that F6 was the most Fe-oxide rich, with
4 perhaps a small secondary peak in F2 (Table 1). XRD (Fig. 6) confirmed visual evidence that
5 magnetically separable material in F6 was almost exclusively magnetite with only small amounts
6 of albite, quartz, and hematite.

7 8 *3.3. Organic matter composition* 9

10 Lignin, as well as cutin and suberin acids, decreased with increasing organo-mineral particle
11 density (Table 1), both decreasing about 75% from F1 to F6. Lignin oxidation for vanillyl
12 phenols increased monotonically with increasing density, whereas the oxidation index for
13 syringyl phenols showed much less pattern. For most fractions the ratio of cutin and suberin
14 acids to lignin ranged between 2 and 3 except for F3 (ratio ~1). Lignin S/V and C/V ratios
15 varied little with density. The values indicate, based upon published values, that gymnosperm
16 needles contribute more lignin to this soil than woody tissue (Hedges and Mann, 1979). They
17 also indicate little contribution by the abundant understory and herbaceous angiosperms, except
18 perhaps to F5 and F6.

19 20 *3.4 Isotope patterns* 21

22 Radiocarbon values ($\Delta^{14}\text{C}$) increased with density, then decreased across F3 to F6, reflecting a
23 decreasing incorporation of “bomb carbon” from atmospheric nuclear weapons testing in the
24 1960s as well as high charcoal content in F1 and F2. It is likely that much of this charcoal
25 originated from a stand-replacement fire that occurred approximately 500 years ago (Grier and
26 Logan, 1977). The MRTs of F1 and F2 were not calculated, pending isolation and radiocarbon
27 analysis of the charcoal. The MRTs of F3-F6 increased steadily with density. Both ^{13}C and ^{15}N
28 values increased with density across F1 to F3 (-26.7 to -24.5 ‰, and 0.03 to 5.42 ‰), then
29 leveled off.

30 31 32 *3.5 Calculated mineral phase densities and organic- vs. mineral-phase volume* 33

34 We calculated the range in density of the mineral phase (Table 1) from the % C and end-point
35 densities of each fraction, assuming that: (1) the density of the organic phase was 1.4 g cm^{-3} (see
36 Adams, 1973; Mayer et al., in press) and (2) the organic phase of all fractions contained 50% C.
37 For the five organo-mineral fractions (F2-F6), the calculated mineral-phase density varied
38 erratically between 2.1 and 2.6 g cm^{-3} (Table 1), typical of many primary and secondary soil
39 minerals (Klein and Hurlbut, 1993). Values for all five fractions were higher, however, than the
40 theoretical value for allophane (1.9 g cm^{-3}), consistent with the XRD results showing little
41 allophane. The especially high range for F2 is consistent with the slightly higher level of
42 magnetically separable material in that fraction.

43
44 To explore whether heavier particles imply thinner organic layers, we calculated the average
45 volume of the organic phase (layer) relative to the total organo-mineral particle for each density
46 fraction, assuming again a 50% C concentration and a density of 1.4 g cm^{-3} for the organic phase.

1 Percent organic volume decreased with increasing density from 81 to 2% (Table 1), as it must
2 given the decreased % C and our assumptions. The decreased ratio of organic phase to total
3 volume is consistent with a corresponding decrease in layer thickness. A better estimate of layer
4 thickness, however, will require data on particle size and both mineral- and organic-phase
5 density.
6
7

8 **4. Discussion**

9

10 *4.1 Comparison with previous studies*

11

12 Our results for OM composition across the density fractions are generally consistent with
13 previously reported results. The peak in the dry-weight distribution of this soil in F4 (Fig. 4) is
14 striking but consistent with the low A-horizon bulk density (0.82 g cm^{-3}) typical of the DIRT
15 study-plot soils (Sulzman et al., in press). C and N concentrations fit well with reports cited in
16 Figs. 1 and 2. C/N of F1 and F2 was higher than those reported elsewhere (Fig. 1), probably due
17 to the large amounts of charcoal in our soil sample. C/N, however, decreased more across the
18 three densest fractions in previous studies than in ours.
19

20 The decrease in lignin phenols with increasing density is consistent with previous reports that
21 lignin is not an especially long-lived component of SOM (Nierop, 2001; Nierop and Verstraten,
22 2003; Rumpel et al., 2004). The lignin phenols that do persist are in an increasingly oxidized
23 state (Table 1), which permits strong sorption to hydroxylated mineral surfaces by ligand
24 exchange. The relatively constant ratio of cutin and suberin acids to lignin phenols, however,
25 indicates that the two classes of compounds are equally distributed across the density series. The
26 ratio of syringyl to vanillyl (S/V) and cinnamyl to vanillyl (C/V) compounds for each of the
27 fractions is within the range previously reported for gymnosperm wood and needles (Hedges and
28 Mann, 1979; Opsahl and Benner, 1995).
29

30 Baisden et al. (2002) provided the only other published isotope data for SDFs of which we are
31 aware. Some of their soils yielded decreases in $\Delta^{14}\text{C}$ with increasing density but others showed
32 increases or held constant. Their ^{13}C values generally increased with particle density, but
33 interpretation of their results is complicated by their use of sonication before each increase in
34 density, which may have redistributed organics into the higher-density fractions. Also, for
35 several of their soils, their densest fraction ($> 2.22 \text{ g cm}^{-3}$) accounted for virtually all the soil C.
36 This result, at variance with most of the SDF studies cited in Fig. 1, could be a result of unusual
37 soil mineralogy or texture or perhaps an artifact of the repeated sonication.
38

39 *4.2 The layer-thickness and “onion” layering hypotheses.*

40

41 Given the large differences in mineralogy of the fractions it was difficult to separate effects of
42 OM content on organo-mineral particle density from those of mineral-phase density. Our
43 calculations (Table 1) suggest that variation in mineral-phase density is unlikely to explain all the
44 variation in organo-mineral particle density, whereas %C correlated well with particle density, as
45 it has in previous studies (Fig. 2). Our data are also consistent with a decrease in organic layer
46 thickness with increasing particle density, although a definitive answer will require additional

1 data. SDF of soils of simpler mineralogy (smectite dominated Alfisol and Fe-oxide dominated
2 Oxisol) is underway and may provide more readily interpretable results.

3
4 The increase in oxidation index (Ac/Al) for vanillyl (though not syringyl) phenols is consistent
5 with our “onion” layering model, in which carboxylated and peptidic compounds sorb
6 preferentially as a stable inner layer directly on mineral surfaces. The increase in ^{14}C MRT is
7 also consistent, as is the monotonic decrease in C/N with increasing density seen for many soils
8 (Fig. 1). The relatively consistent C/N across F4-F6 of our forest soil suggests a common organic
9 constituent. Surface-active microbial proteins are an obvious possibility. The role of such
10 proteins in bacterial attachment has been extensively studied (Marshall, 1971; Bitton and
11 Marchall, 1980). Cellular attachment is preceded by protein sorption to mineral surfaces (Bashan
12 and Levanony, 1988; Dufrene et al., 1996, 1999). Mycelial fungi secrete a broad class of
13 surface-active glycoproteins, the hydrophobins, with discrete hydrophilic and hydrophobic
14 domains, allowing formation of amphipathic layers (Wösten, 2001; Wessels, 1997). The ability
15 of hydrophobins to bind to either hydrophilic or hydrophobic surfaces would permit sorption
16 across a range of mineral surfaces. Research into the role of hydrophobins in SOM dynamics has
17 been restricted to one group attributed to arbuscular mycorrhizal fungi, the glomalins, suggested
18 to be a significant component of SOM and agents of aggregate formation (Wright et al., 1999;
19 Franzluebbers et al., 2000; Rillig et al., 2002). The ability of hydrophobins to bind to either
20 hydrophilic or hydrophobic surfaces would permit sorption across a range of mineral surfaces
21 and thus explain to some extent the consistency in organic matter C/N across the different types
22 of minerals observed in the three heaviest density fractions. A clearer indication of the relative
23 roles of cation-exchanged peptidic compounds and ligand-exchanged carboxyl compounds in
24 forming the inner layer in an “onion” model requires quantification of these various compounds.

25 26 27 *4.3 Organic matter composition in relation to mineralogy*

28
29 Sequential density fractionation (SDF) separated the soil particles into three major groupings. F1
30 and F2 were mainly plant debris, F3 and F4 accounted for most of the layer silicate clays in this
31 soil, and F5 and F6 contained mainly primary minerals. Differences in OM composition across
32 the fractions must be interpreted in the context of these mineralogical differences.

33
34 Most measures of OM composition (Table 1) showed one of two patterns: A, a monotonic
35 change with increasing density, or B, a monotonic change across F1 through F4, then little
36 change across F5 and F6. For example, total C, N, lignin phenols, vanillyl lignin oxidation
37 index, and cutin+suberin acids all showed pattern A, whereas C/N, ^{13}C , ^{15}N , and syringyl lignin
38 oxidation index showed pattern B. Estimates of ^{14}C MRT showed pattern A over the four
39 organo-mineral fractions (F3-F6).

40
41 The turnover time for the bulk of the OM increased with increasing organo-mineral particle
42 density in this soil, indicating more stable sorption of the organics with increasing particle
43 density. Because ligand exchange is the most stable of the various sorption mechanisms (Sposito,
44 1988), the MRTs suggest that ligand exchange becomes more prominent with increasing density,
45 and presumably with decreasing layer thickness. Conversely, the less dense organo-mineral
46 fractions may include more recently stabilized C in the outer layers of thicker coatings. Thus,

1 although the inner layers of these lower density particles may have MRTs similar to those of
2 heavier particles, the average MRT is lower. The increase in lignin oxidation with density is
3 consistent with this pattern, as increased oxidation implies more carboxyl groups available to
4 ligand exchange with the variable-charge (hydroxylated) mineral surfaces. The decrease in C/N
5 with density reported for most soils worldwide, though only across the lightest four fractions for
6 this soil, suggests in addition a role for electrostatically-bound peptidic compounds in stable
7 sorption of the inner layers and thus in slow OM turnover.

8
9 The ^{14}C MRTs must, however, be considered in the context of the stable isotope data reported
10 here. There is ample evidence that both ^{13}C and ^{15}N accumulate preferentially as a result of
11 continued microbial processing, though the mechanisms are still unclear (Gleixner et al., 1993;
12 Hayes 2001; Filley et al., 2001; Abraham and Hesse, 2003; Fernandez et al., 2003). The lack of
13 change in ^{13}C and ^{15}N across the three heaviest fractions might be explained by especially stable
14 sorption (as evidenced by the older C MRTs) with consequently reduced opportunity for further
15 microbial processing and isotope fractionation. In any case, any explanation of isotope
16 fractionation based on microbial processing must explain how enzymes can make contact with,
17 and then cleave, organics that are sorbed to mineral surfaces, especially in the inner layers,
18 access to which by both organisms and enzymes is presumably hindered by overlying organics.

19
20 An alternative, though not mutually exclusive, explanation is that organics in the heavier
21 particles were already rich in ^{13}C and ^{15}N at the time that they sorbed to the mineral surfaces.
22 Lignin as well as cutin and suberin acids, all of which decreased with increasing particle density
23 in our andic soil, are well known to be depleted in ^{13}C relative to bulk plant material (Deines,
24 1980). Other compounds, especially amino acids, may be enriched or depleted in both ^{13}C and
25 ^{15}N depending on the specific compound and the substrate on which the organisms that produce
26 the compounds are grown (Macko et al., 1987; Hayes, 2001; Abraham and Hesse, 2003). More
27 information, especially compound-specific ^{15}N data, will be needed to explore fully this
28 explanation.

29 30 31 **5. Concluding Remarks**

32
33 Removal of particulate organic debris (particle density $<1.85\text{ g cm}^{-3}$) allowed focus on the
34 mineral-dominated fractions. Subsequent isolation of several sequential density fractions, rather
35 than a single “heavy fraction”, then permitted more ready interpretation of the results of ^{14}C , ^{13}C ,
36 ^{15}N , lignin, and other analyses. The sequential fractionation also isolated discrete mineralogies,
37 allowing exploration of relations between mineral surface chemistry and composition of the
38 associated organic matter.

39
40 Our data are consistent with an “onion” model of SOM accumulation on mineral surfaces in
41 which polar carboxyl and nitrogenous compounds sorb by ligand and cation exchange. Such
42 organics then provide a relatively nonpolar surface onto which hydrophobic and less polar
43 hydrophilic organics can sorb more stably than they could onto clean mineral surfaces. An
44 “onion” layering model begs questions about mechanisms of renewal and production of clean
45 mineral surfaces in soils. This could occur by breakage of existing primary or secondary
46 minerals, and by aeolian or lateral inputs of new mineral material. Other mechanisms, such as

1 incongruent dissolution of primary minerals and subsequent formation of secondary minerals, as
2 well as passage through soil animal guts, may also be important.

3
4 Our results also imply that sorption of peptidic organics to soil mineral particles is a potentially
5 important process of C and N stabilization and deserves further study. To the extent that the N-
6 containing organics sorb mainly electrostatically, and the non-N-containing organics sorb mainly
7 via ligand exchange and weaker interactions, nitrogenous and non-nitrogenous compounds are
8 stabilized via different mechanisms. This distinction is not reflected in current SOM models of
9 SOM dynamics (Jenkinson et al., 1991; Parton et al. 1991; McGuire et al., 1997; Currie, 2003;
10 Rastetter et al., 2005). A better understanding of the processes that both stabilize and destabilize
11 SOM would improve our ability to predict future trends in agricultural and forest productivity
12 and C exchange between soil and atmosphere.

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15
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Table 1
Measured and calculated properties of the 6 density fractions.

	Fraction density (g cm ⁻³)					
	F1	F2	F3	F4	F5	F6
	<1.65	1.65-1.85	1.85-2.00	2.00-2.28	2.28-2.55	>2.55
Dry mass (% of soil)	2.98	1.84	10.7	57.5	16.9	10.1
pH in NaF	8.7	9.1	9.6	9.9	10.0	9.8
Magnetic mass (% of fraction)	0.07	0.29	0.20	0.13	0.08	10.5
Total C (% of fraction)	36.0	28.5	14.1	2.69	0.74	0.22
Total N (% of fraction)	0.54	0.55	0.46	0.23	0.062	0.019
C:N ratio	66	52	30	12	12	12
$\Delta^{14}\text{C}$ concentration (‰)	6	36	55	30	-59	-95
¹⁴ C mean residence time (y)	ND	ND	150	210	680	985
$\delta^{13}\text{C}$ (‰)	-26.67	-26.21	-25.93	-24.50	-24.73	-24.66
$\delta^{15}\text{N}$ (‰)	0.03	2.67	4.19	4.80	5.42	4.83
Total lignin yield (mg g ⁻¹ C)	31	25	19	16	8.4	7.8
Lignin oxidation index (vanillyl) ^a	0.49	0.71	0.92	1.31	1.38	1.48
Lignin oxidation index (syringyl) ^a	0.51	0.48	0.72	0.97	0.70	0.76
Cutin+suberin acids (mg g ⁻¹ C)	66	77	58	19	14	16
C/V (S/V) ^b	0.17 (0.02)	0.29 (0.02)	0.25 (0.04)	0.28 (0.07)	0.22 (0.23)	0.25 (0.31)
Calculated mineral-phase density range (g cm ⁻³)	n.a.	2.2-3.2	2.1-2.4	2.1-2.3	2.3-2.6	>2.6
Volume of organic phase relative to total particle volume (%)	100	58	29	5.5	1.5	0.4

^aRatio of carboxylic acid to aldehyde functional groups for vanillyl lignin phenols.

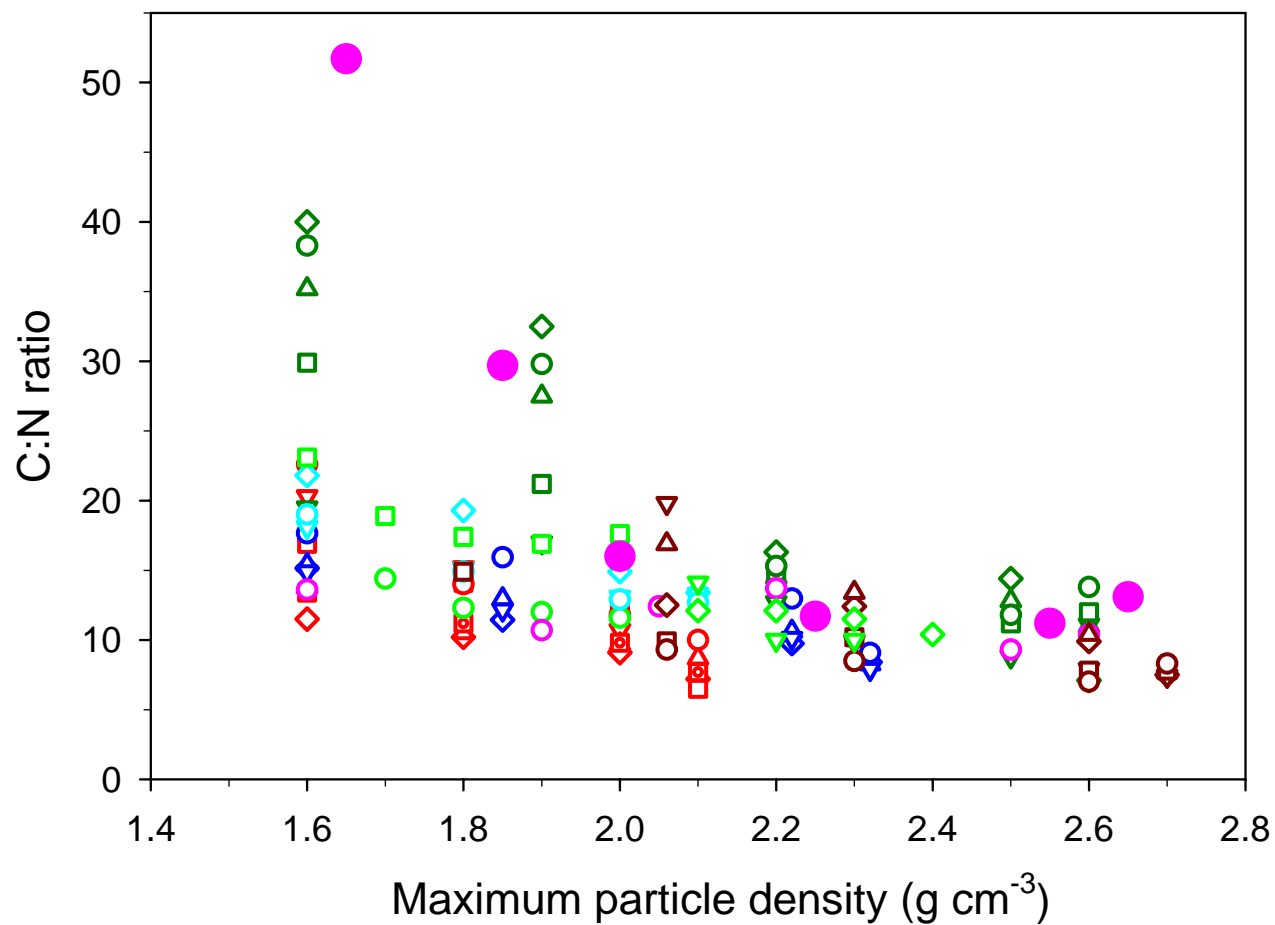
^bRatio of the sum of cinnamyl (or syringyl) phenols to vanillyl phenols.

n.a. -- not applicable

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8 two fractions.
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20 Fig. 6. X-ray diffraction traces for magnetic and non-magnetic portions of bulk soil and for the
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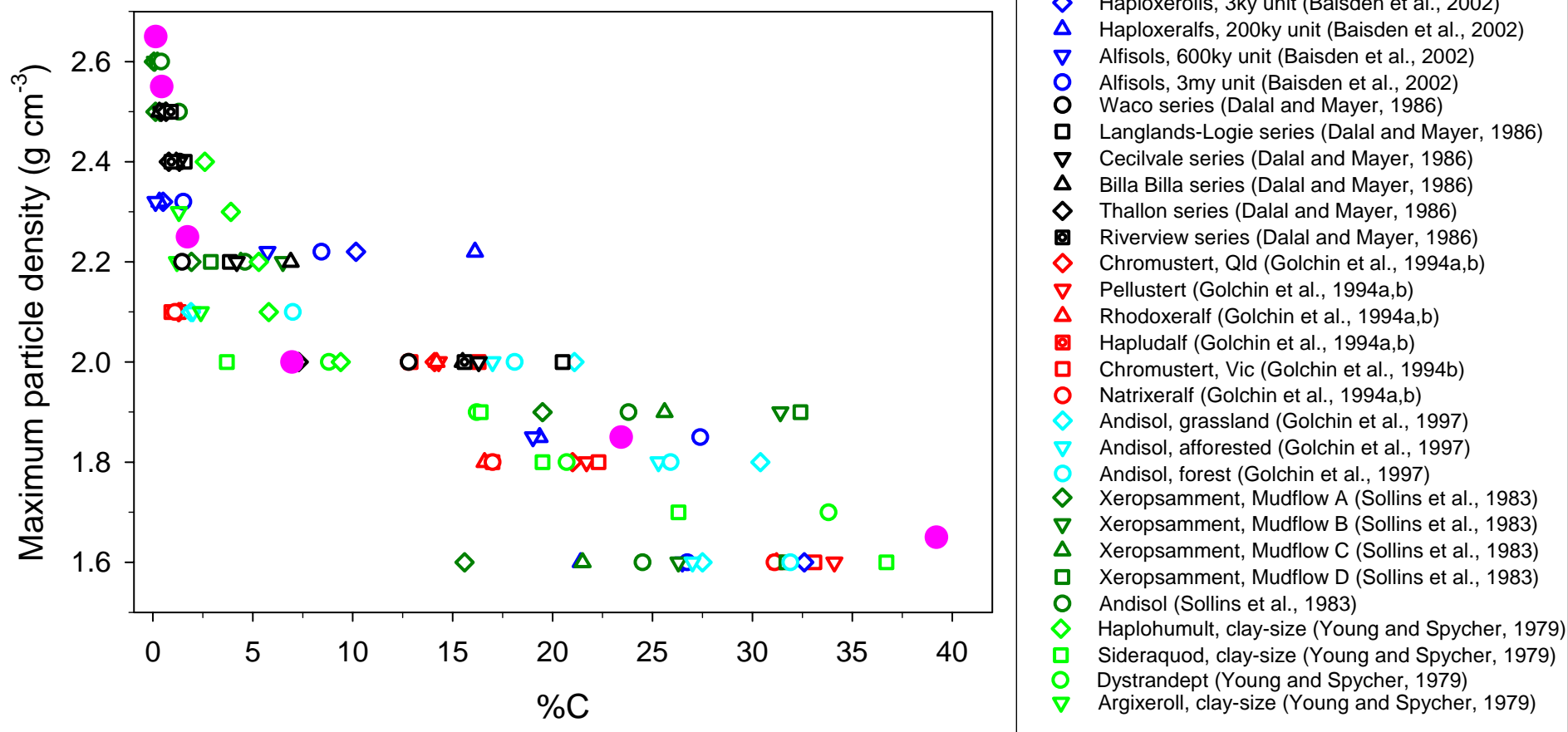


- Andic Dystrudept (this study)
- Marine sediment (Arnarson and Keil, 2001)
- ◇ Haploxerolls, 3ky unit (Baisden et al., 2002)
- △ Haploxeralfs, 200ky unit (Baisden et al., 2002)
- ▽ Alfisols, 600ky unit (Baisden et al., 2002)
- Alfisols, 3my unit (Baisden et al., 2002)
- ◇ Chromustert, Qld (Golchin et al., 1994a,b)
- ▽ Pellustert (Golchin et al., 1994a,b)
- △ Rhodoxeralf (Golchin et al., 1994a,b)
- ▣ Hapludalf (Golchin et al., 1994a,b)
- ▣ Chromustert, Vic (Golchin et al., 1994b)
- Natrixeralf (Golchin et al., 1994a,b)
- ◇ Andisol, grassland (Golchin et al., 1997)
- ▽ Andisol, afforested (Golchin et al., 1997)
- Andisol, forest (Golchin et al., 1997)
- ◇ Xeropsamment, Mudflow A (Sollins et al., 1983)
- ▽ Xeropsamment, Mudflow B (Sollins et al., 1983)
- △ Xeropsamment, Mudflow C (Sollins et al., 1983)
- ▣ Xeropsamment, Mudflow D (Sollins et al., 1983)
- Andisol (Sollins et al., 1983)
- ◇ Haploboroll, >200µm (Turchenek and Oades, 1979)
- ▽ Haploboroll, 20-5µm (Turchenek and Oades, 1979)
- △ Haploboroll, 5-2µm (Turchenek and Oades, 1979)
- ▣ Haploboroll, 2-0.4µm (Turchenek and Oades, 1979)
- Haploboroll, <0.4µm (Turchenek and Oades, 1979)
- ◇ Haplohumult, clay-size (Young and Spycher, 1979)
- ▣ Sideraquod, clay-size (Young and Spycher, 1979)
- Dystrandept (Young and Spycher, 1979)
- ▽ Argixeroll, clay-size (Young and Spycher, 1979)

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Fig. 1. C:N ratio vs. particle density for soils worldwide.

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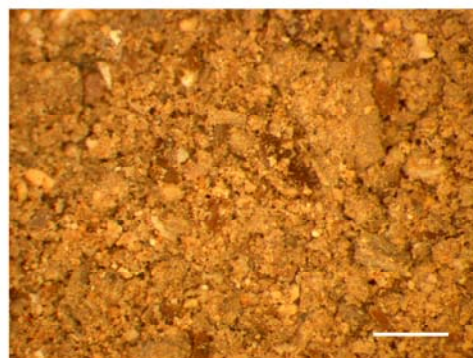


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Fig. 2. Particle density vs. %C for soils worldwide.



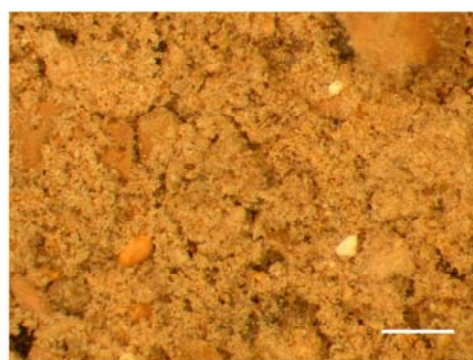
$< 1.65 \text{ g cm}^{-3}$



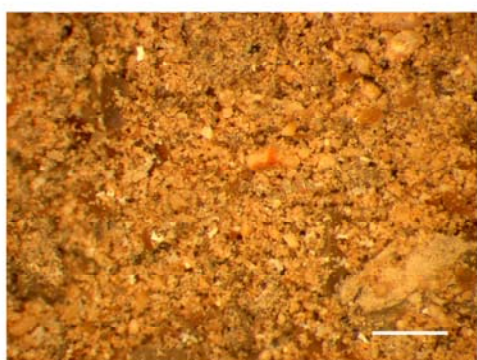
2.00 - 2.28



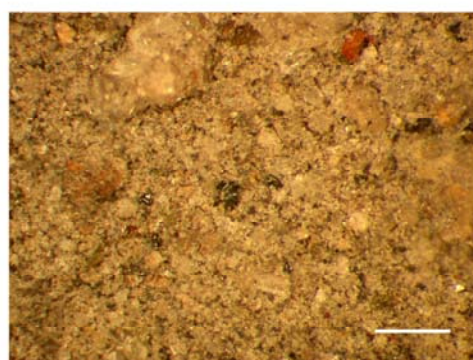
1.65 - 1.85



2.28 - 2.60



1.85 - 2.00



> 2.60

Fig. 3. Light microscope images of the six density fractions.

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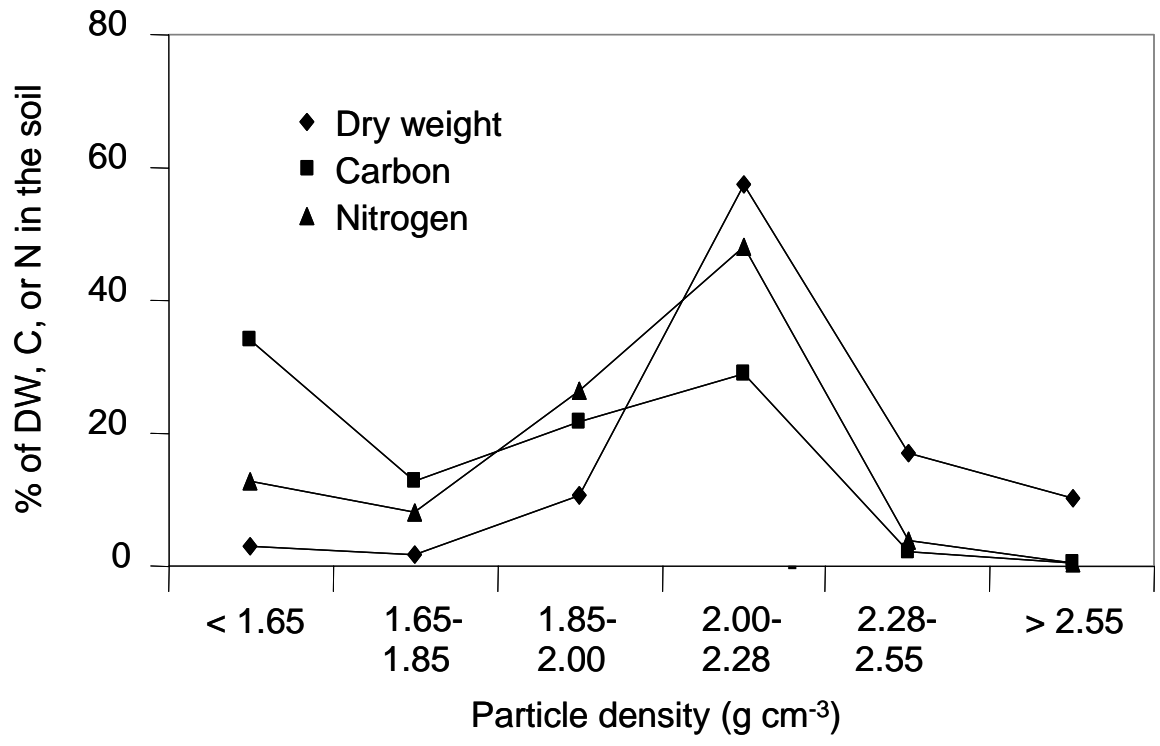
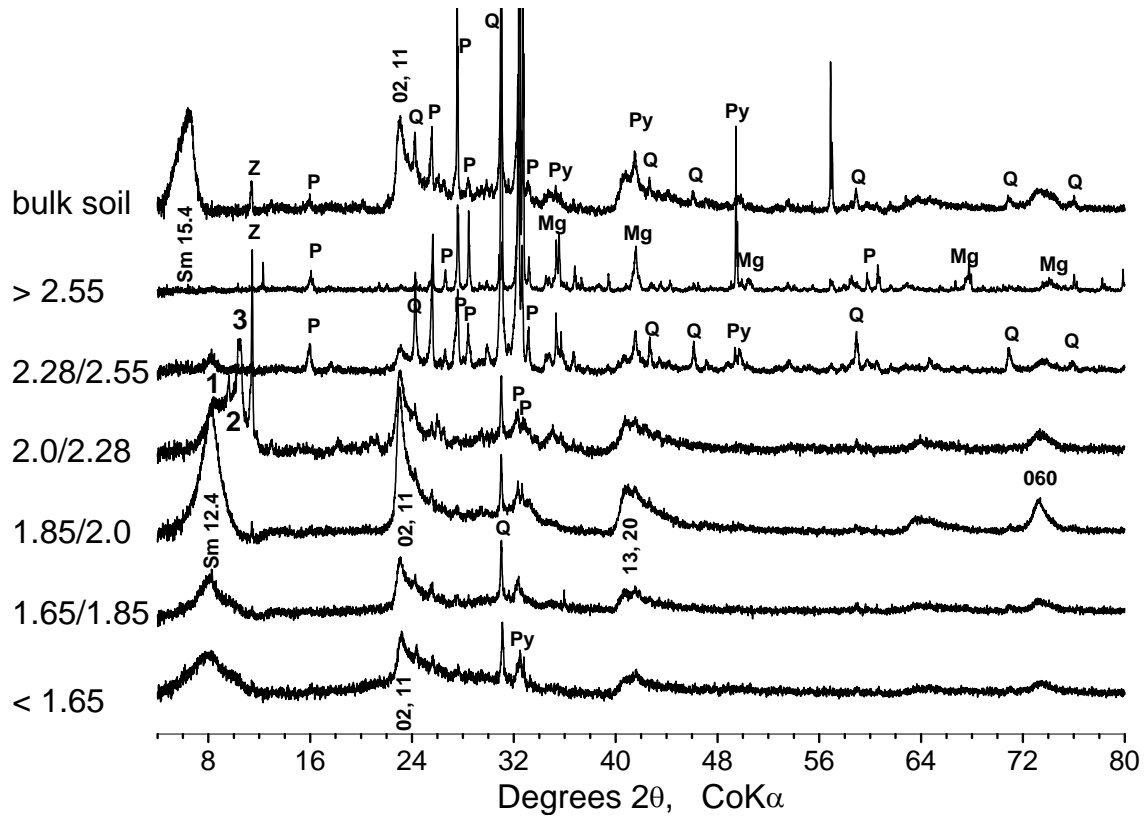


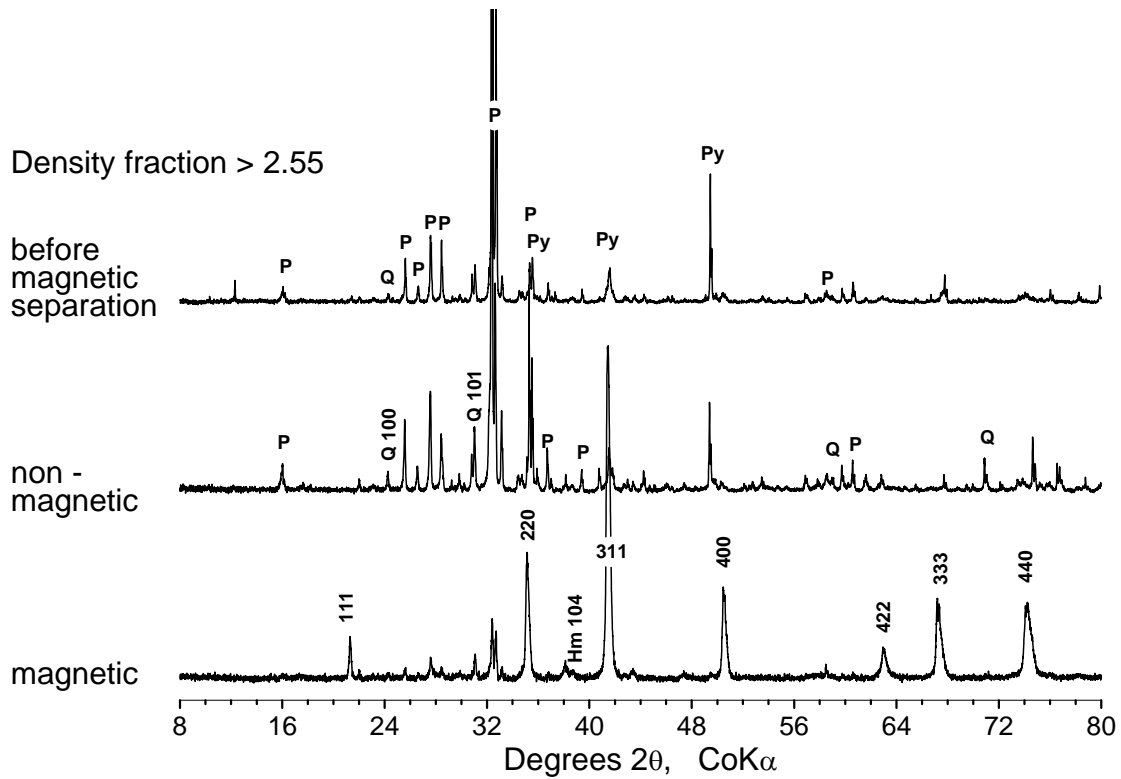
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Fig. 6. X-ray diffraction traces for magnetic and non-magnetic portions of bulk soil and for the heaviest density fraction ($>2.55 \text{ g cm}^{-3}$). P = plagioclase feldspar, Q = quartz, Py = pyroxene, Hm = hematite. Q 100: Q 101 and Hm 104 are hkl reflections of quartz and hematite, respectively. Unlabeled numbers are hkl reflections of magnetite.