

## **Test Results for CSTR Test 4**

**D. D. Lee  
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Chemical Technology Division

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## ACRONYMS

ALARA	as low as reasonably achievable
BDL	below the detection limit
CERS	Chemical and Energy Research Section
CsTPB	cesium tetraphenylborate
CSTR	continuous-flow stirred-tank reactor
DF	decontamination factor
FTIR	Fourier transform infrared analyzer
HLW	high-level waste
HPLC	high-performance liquid chromatography
ICP	inductively coupled plasma
ICP-MS	inductively coupled plasma–mass spectroscopy
KTPB	potassium tetraphenylborate
MST	monosodium titanate
NaTPB	sodium tetraphenylborate
NIST	National Institute of Standards and Technology
ORNL	Oak Ridge National Laboratory
PBs	1PB, 2PB, and 3PB decomposition products of TPB
P&E	Plant and Equipment Division (ORNL)
QA	quality assurance
RMAL	Radioactive Materials Analytical Laboratory
SP	sample port
SPP	Salt Waste Processing Program
SRL	Savannah River Laboratory
SRS	Savannah River Site
SRTC	Savannah River Technology Center
STTP	Small-Tank Tetraphenylborate Process
TPB	tetraphenylborate
TTP	technical task plan
WAC	waste acceptance criteria
WSRC	Westinghouse Savannah River Company
3PB	triphenylborane
2PB	diphenylborinic acid
1PB	phenylboronic acid



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## ABSTRACT

The goal of the Savannah River Salt Waste Processing Program (SPP) is to evaluate and select the most effective technology for the treatment of high-level waste salt solutions presently being stored in underground storage tanks at the United States Department of Energy Savannah River Site (SRS) in Aiken, South Carolina. One of the three technologies currently being developed as part of this effort is the Small-Tank Tetraphenylborate Process (STTP). This process uses sodium tetraphenylborate (NaTPB) to precipitate and remove radioactive cesium from the waste and monosodium titanate (MST) to sorb and remove radioactive strontium and actinides. Oak Ridge National Laboratory is demonstrating this process at the 1:4000 scale using a 20-L-capacity continuous-flow stirred tank reactor (CSTR) system. Since March 1999, four operating campaigns of the 20-L CSTR have been conducted. The ultimate goal is to verify that this process, under certain extremes of operating conditions, can meet the minimum treatment criteria necessary for processing and disposal at the Savannah River Saltstone Facility. The waste acceptance criteria (WAC) for  $^{137}\text{Cs}$ ,  $^{90}\text{Sr}$ , and total actinides are  $<40$  nCi/g,  $<40$  nCi/g, and  $<18$  nCi/g, respectively. However, to allow for changes in process conditions, SPP is seeking a level of treatment that is about 50% of the WAC. The bounding separation goals for  $^{137}\text{Cs}$  and  $^{90}\text{Sr}$  are to obtain decontamination factors (DFs) of 40,000 (99.998% removal) and 26 (96.15% removal), respectively. The DF is defined as the concentration of contaminant in the waste divided by the concentration of contaminant in the effluent stream.

Tests 1 and 2 were conducted in June and July 1999, and the results are documented in ORNL/TM-1999/234. These initial tests, using simulants traced with radioactive materials, verified that the STTP process could achieve the necessary cesium, strontium, and actinide decontamination under standard operating conditions, with and without the recovery and recycle of the excess TPB that was not precipitated with potassium and cesium. TPB decomposition, the major side reaction of the process, did not occur in Test 2 despite the addition of materials [SRS synthetic sludge and modified Enhanced Comprehensive Catalyst (ECC)] that were known to cause TPB to degrade. Test 3 was a 72-h demonstration that was conducted during August 2000. The objectives were to examine the effectiveness of an improved foam suppressant (IIT B52) in minimizing foam formation and to determine if its presence had a deleterious effect on the removal of Cs, Sr, and U from simulated waste. Test 3 was operated long enough to provide about nine system turnovers. The DFs for Cs, Sr, and U obtained for the filtrate from the Slurry Concentrating Tank were  $>40,000$ ,  $\sim 50$ , and  $\sim 5$ , respectively, which exceeded the WAC standards. No foaming problems were experienced during the test, and the foam suppressant had no apparent negative effect on the separations process. High-performance liquid chromatography (HPLC) analyses indicated that no measurable NaTPB decomposition occurred during the test. The results of Test 3 are documented in CERS/SR/TPB/011.

Test 4 was a 130.3-h demonstration that was conducted in October 2000. As in Test 3, foam suppressant IIT B52 was added to the system vessels to minimize foam formation. The primary goal was to verify that the STTP process could achieve and maintain the necessary cesium decontamination while TPB was actively decomposing. An improved SRS catalyst system, consisting of Pd(0) on alumina powder, Hg(II) salt, phenylboronic acid (1PB), and benzene, was continuously added during the test, which successfully initiated and sustained TPB decomposition. The rate of TPB decomposition was determined by analyzing process samples for TPB degradation products using reverse-phase HPLC. The rate of decomposition was also sensitive to temperature.

The system vessels were controlled at 25°C during the first 76 h of operation. The temperature was subsequently increased to 45°C over the next 13 h and then maintained at that temperature for the remainder of the test. As expected, the TPB decomposition rate increased at the elevated temperatures. At the end of the test, about 4.2% of the total TPB in the first CSTR vessel was being decomposed. This behavior was also confirmed by on-line Fourier transform infrared (analyzer) measurement of benzene from the vapor space of each vessel. Even with TPB being decomposed by the off-normal conditions of this test, the DF for <sup>137</sup>Cs obtained for the filtrate from the Slurry Concentrating Tank ranged from 47,000 to 646,000, exceeding the WAC standard.

Overall, the system control and hydraulic behavior were good, with no foam formation or problems being detected in any of the vessels. More than 162 L of radioactive, simulated waste was treated during the test, and ~218 L of decontaminated filtrate was collected. About 13 L of 10 wt % TPB precipitate/MST slurry was collected and washed to recover the excess TPB (that was not precipitated as potassium and cesium). Only about 10 to 11% of the available TPB was recovered during slurry washing with water.



# 1. INTRODUCTION

## 1.1 BACKGROUND

The Small-Tank Tetraphenylborate Process (STTP) is one of three separations processes currently being examined by the Savannah River High-Level Waste (HLW) Salt Waste Processing Program (SPP) as an alternative to the In-Tank Precipitation Process for treatment of highly radioactive Savannah River Site (SRS) tank waste. The ultimate goal of the process is to decontaminate the SRS salt waste so that it meets the waste acceptance criteria (WAC) for processing and disposal at the Savannah River Saltstone Facility. The WAC for  $^{137}\text{Cs}$ ,  $^{90}\text{Sr}$ , and alpha are  $<40$  nCi/g,  $<40$  nCi/g, and  $<18$  nCi/g, respectively. However, to allow for changes in process conditions, the SPP is seeking a level of treatment that is about 50% of the WAC. Supernatants in "high- $\text{OH}^-$ " SRS tanks that are 6.4 M in  $\text{Na}^+$  have cesium concentrations in the range of 0.000415 M. Cesium concentrations for average SRS waste with the same  $\text{Na}^+$  concentration are in the range of 0.000157 M. The  $^{137}\text{Cs}$  radioactivity levels for these wastes are 1.1 and 0.41 Ci/L, respectively.

A 1:4000-scale, 20-L continuous-flow stirred-tank reactor (CSTR) system was designed and fabricated at Oak Ridge National Laboratory (ORNL) for development of this process. The goal of the 20-L CSTR test program is to evaluate the performance of the 20-L CSTR system under conditions that simulate the anticipated full-scale operation of the system. The basic chemistry of the process makes use of tetraphenylborate (TPB) to remove cesium by precipitation and monosodium titanate (MST) to remove strontium and actinides (Pu, Np, and U) by sorption. The CSTR system was to be operated in a contained and shielded hot cell environment to allow the use of radioactive materials. Testing of the CSTR system would determine if this approach could (1) provide the necessary cesium, strontium, and actinide separations needed for treatment of the SRS salt waste; (2) perform satisfactorily with an improved antifoam agent; and (3) perform satisfactorily while active decomposition of TPB was taking place with and without the recovery and recycle of unreacted TPB.

Test 1a (Tests 1b and 1c were not needed) and Test 2 were conducted in FY 1999 and met performance objectives in verifying that the required separations of cesium, strontium, and actinides could be provided by this process.<sup>1</sup> Test 3 was a 72-h demonstration conducted in FY 2000 to examine the effectiveness of an improved foam suppressant (IITB52) in minimizing foam formation and to determine if its presence had a damaging effect on the removal of Cs, Sr, and U from simulated waste.<sup>2</sup> There was no evidence of excessive foam formation indicated in slurry samples taken or by any operational difficulties. Cesium, strontium, and uranium removal were not significantly impacted by the antifoam, and the DFs were comparable to those obtained in earlier tests. Slurry washing at the end of the test showed no apparent effect of the antifoam on TPB recovery as compared with Tests 1a and 2.

The objective of Test 4 was to demonstrate that acceptable system performance could be maintained while TPB was decomposing. The STTP must remove cesium from the salt to achieve the WAC requirement of 40 nCi  $^{137}\text{Cs}$ /g. This requires a decontamination factor (DF) of  $\sim 18,000$  (99.994% cesium removal) for the 1.1-Ci/L "high- $\text{OH}^-$ " supernatant and a DF of  $\sim 7000$  (99.986% cesium removal) for the 0.41-Ci/L average SRS supernatant. The SPP goal is to achieve a DF of 40,000 (99.998% cesium removal), which would successfully treat supernatants with  $^{137}\text{Cs}$  concentrations up to 2.44 Ci/L. Only a few SRS tanks have concentrations that high.

## 1.2 SYSTEM DESIGN

The apparatus used in the test system was sized and designed to provide reliable data for scaleup to larger systems. All feeds are introduced into CSTR 1, and the degree of decontamination of the process stream in any vessel is always defined based on the feed to this reactor. Most of the precipitation of cesium takes place in CSTR 1. CSTR 2 provides extended residence time for some additional precipitation to occur. The primary role of CSTR 2 in the process is to provide residence time for the MST to sorb the actinides present in the real waste. The test system, which is located in Hot Cell B, Building 4501, includes two CSTRs in series, each with a 15-L working volume, a slurry concentration tank with a cross-flow filter to concentrate the slurry, a concentrated-slurry washing tank with a cross-flow filter, and various tanks and pumps to integrate the operation. Two 55-gal feed tanks with mixing systems were located in Hot Cell C. Figure 1 shows a simplified process flowsheet for the system.

Figures 2–6 provide the updated flowsheets for the TPB process test system in Hot Cell B. Figure 2 shows the CSTR feed systems; Fig. 3, the slurry concentration system; and Fig. 4, the slurry wash system. Figures 5 and 6 show the CSTR benzene monitoring flowsheet. Note that the benzene analyzers are on a platform on the wall adjacent to Hot Cell B in room 126C.

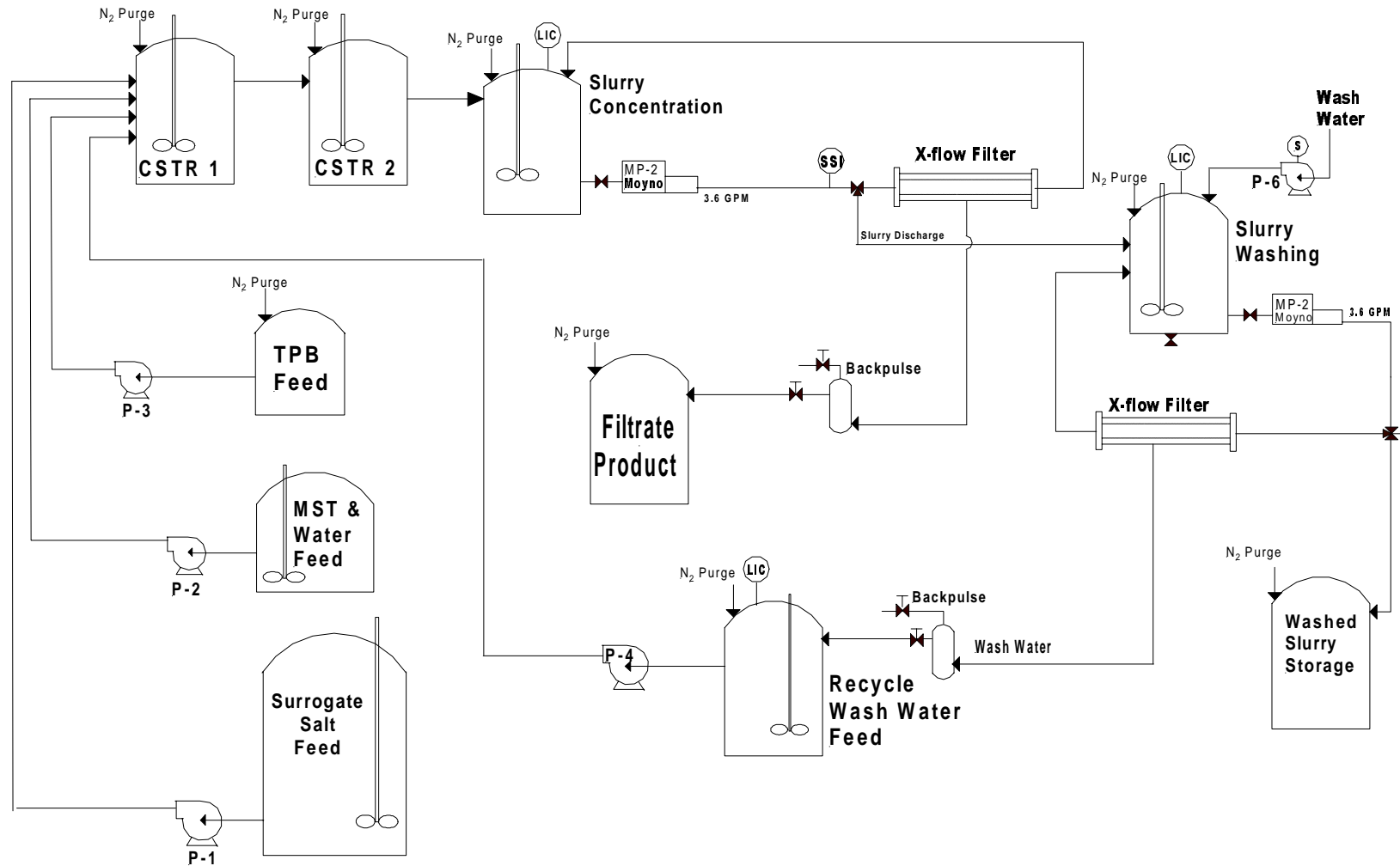


Fig. 1. Simplified diagram of the CSTR systems.

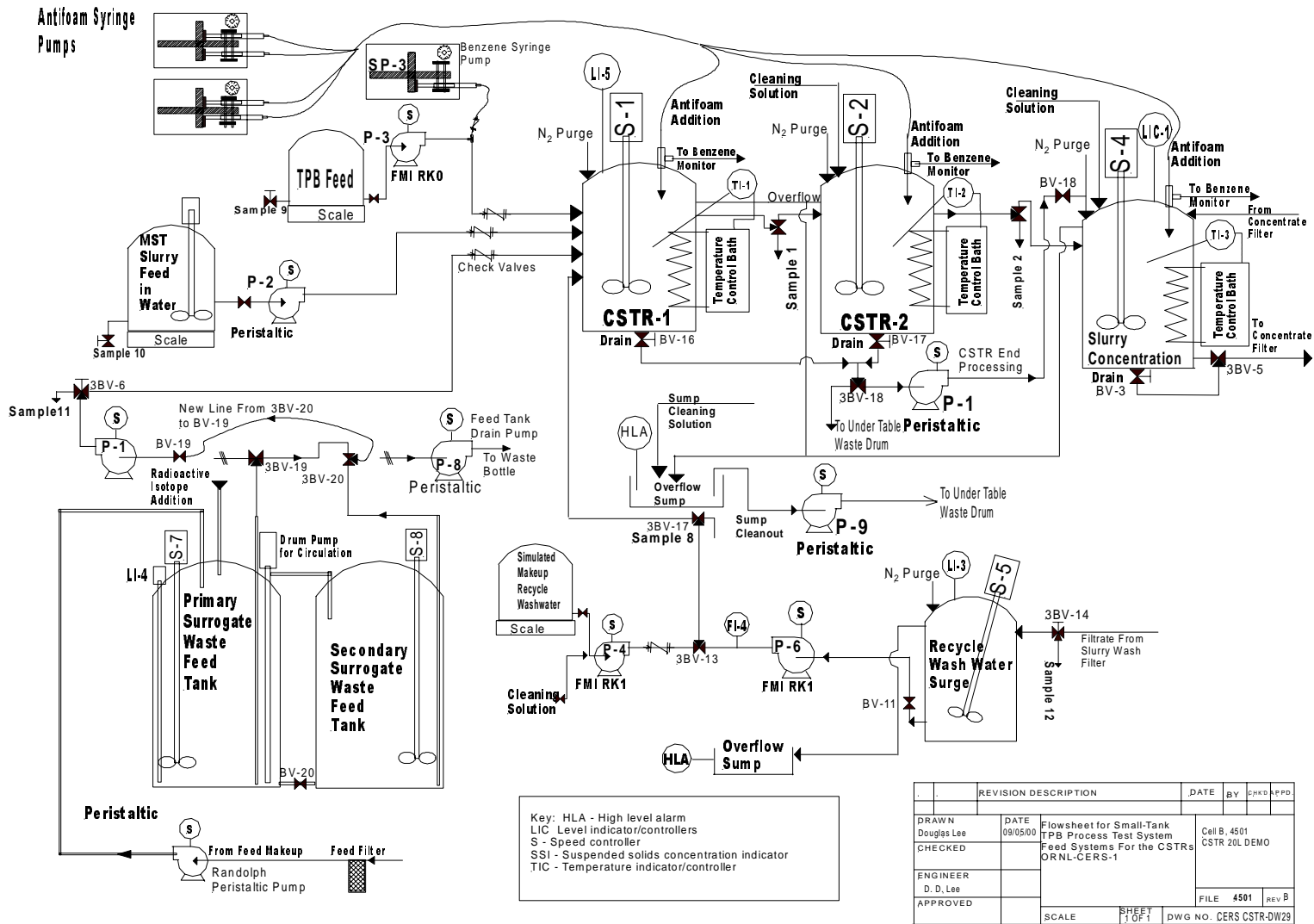


Fig. 2. Diagram of the CSTR feed systems.

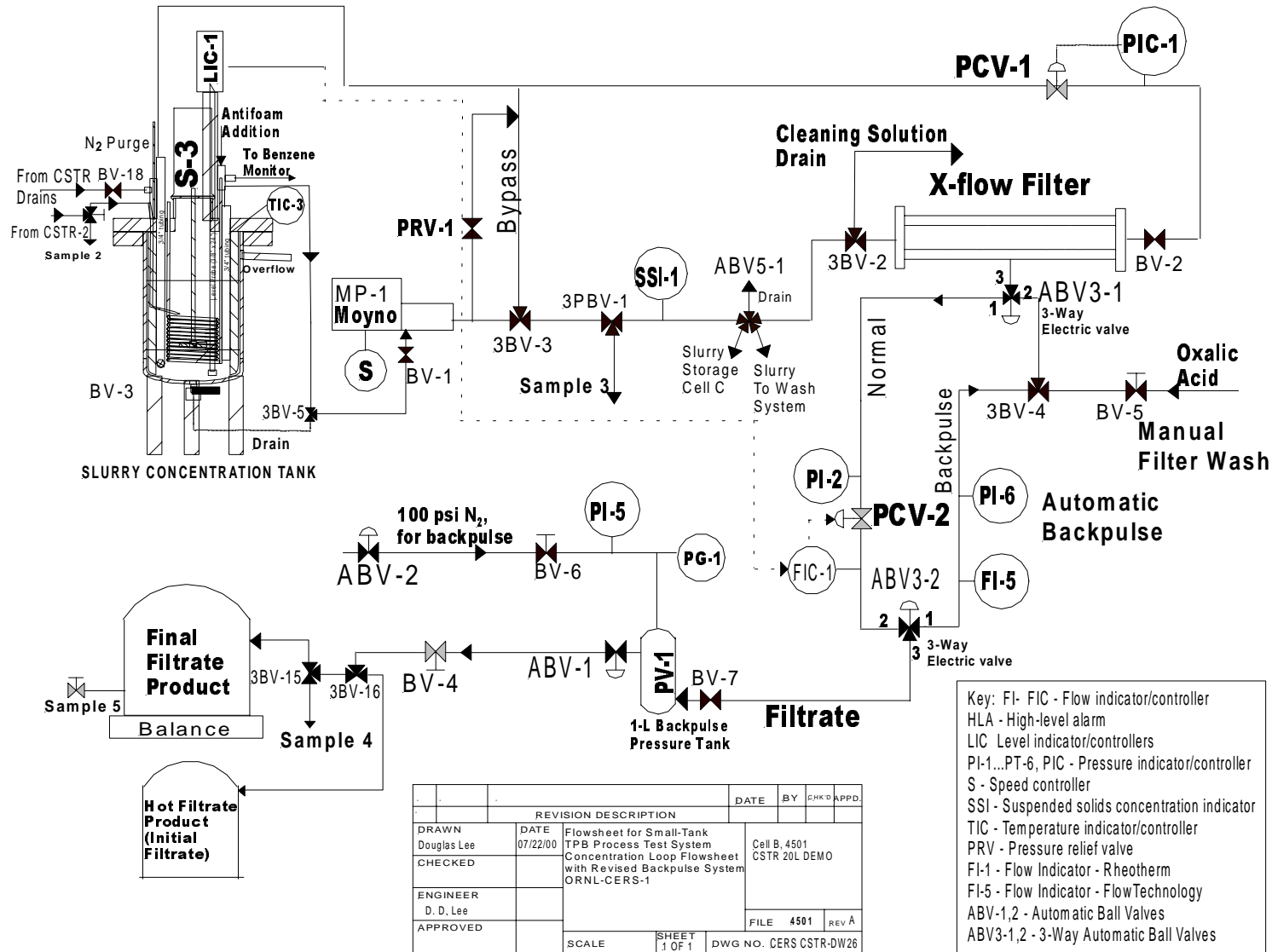


Fig. 3. Redesigned slurry concentration filter system.

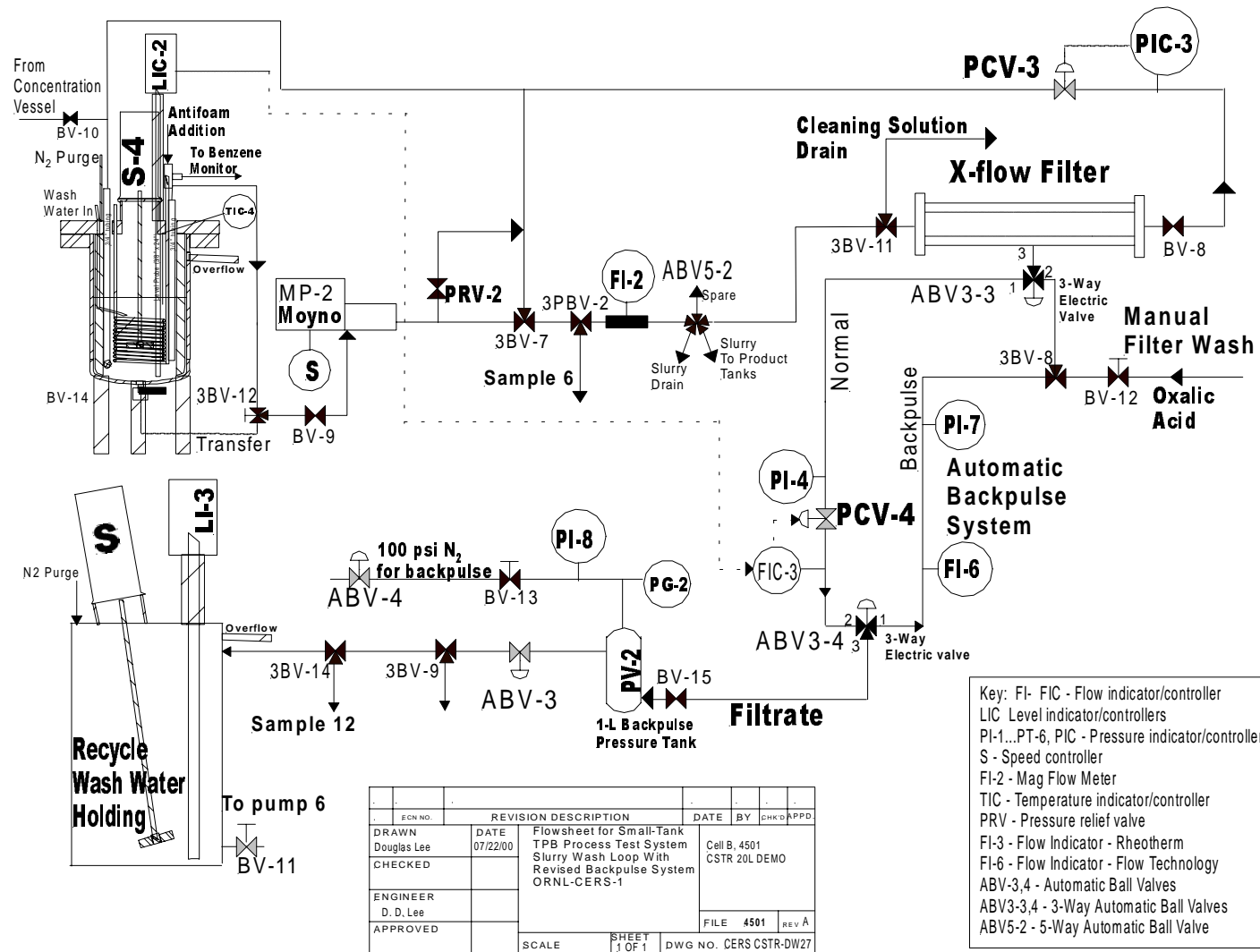


Fig. 4. Redesigned slurry wash filter system.

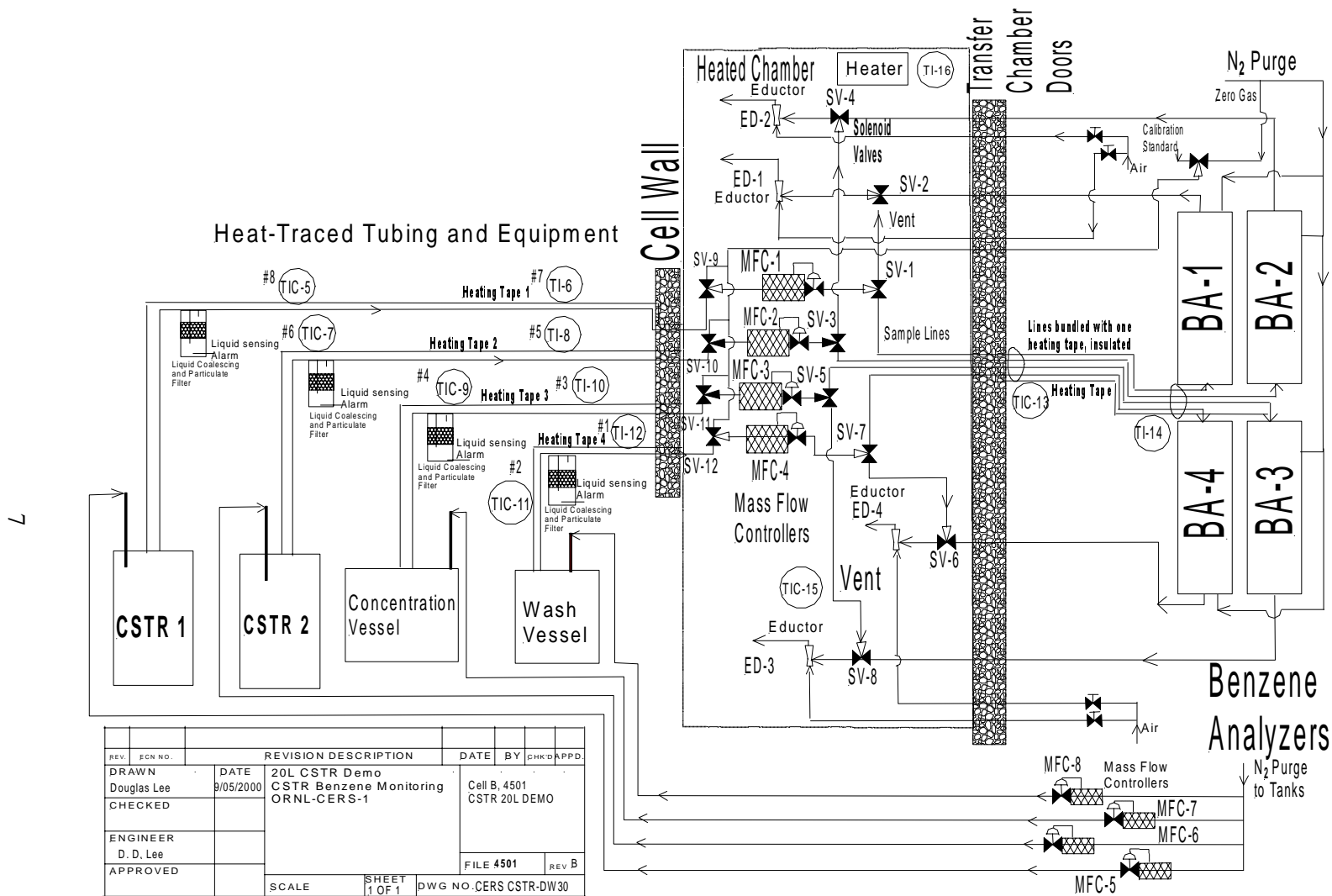


Fig. 5. Flowsheet for the benzene monitoring system.

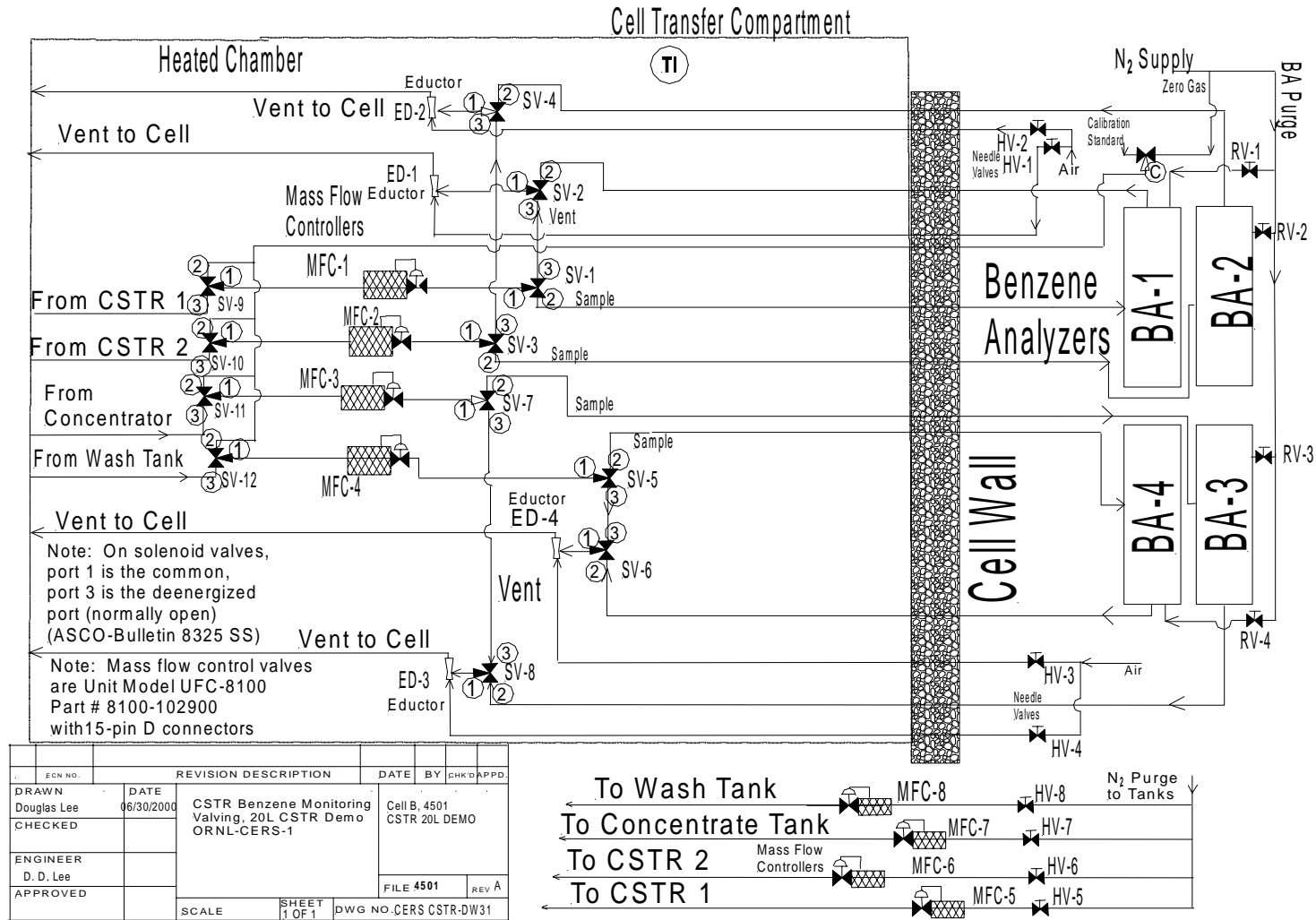


Fig. 6. Valves and piping for the benzene monitoring system.



## 2. TEST 4 PLAN

### 2.1 TEST DESIGN

The primary goal of Test 4 was to evaluate the contaminant removal performance of the CSTR system while TPB was actively decomposing. Test 4 was planned to be a single-pass (no recycle of recovered TPB), 80-h test that provided about ten volume turnovers in the CSTRs. A more aggressive catalyst system was used to cause TPB decomposition and determine if a sustained cesium DF could be maintained under these flowsheet conditions. The catalyst was designed by SRTC (Mark Barnes et al., *Recommendation of Catalyst System and Delivery for ORNL 20-L Testing*, SRT-LWP-2000-00126, August 2000<sup>3</sup>) to provide a “3X” (“X” being the expected catalyst concentration based on historical data) catalyst concentration in the process fluids of the CSTRs. The catalyst consisted of palladium (5% on alumina powder), mercury [ $\text{Hg}(\text{NO}_3)_2 \cdot \text{H}_2\text{O}$ ], phenylboronic acid (1PB), and benzene, which were to be added with the feed materials to CSTR 1. The TPB decomposition goal was to obtain a benzene generation rate in the Slurry Concentrating Tank of about 10 mg/(L·h), which would provide a benzene concentration of about 400 ppm in the purge gas (assuming no holdup of benzene in the process fluids).

A secondary goal of Test 4 was to operate at a higher temperature (45°C) to test system performance under catalytic conditions after the initial goal was reached. The plan involved a 12-h temperature ramp up to 45°C and continued operation at 45°C for the time required to process a second batch of TPB precipitate slurry (10 wt % solids) or until it was clear that the required cesium DF could not be maintained. Test 4 was started with both CSTRs filled with the process fluids remaining in the CSTRs at the end of Test 3 and with the Slurry Concentrating Tank drained and rinsed.

The agitators in the two CSTRs were run at speeds in the range of 650–850 rpm for this test. The concentration of the IITB52 antifoam agent (Batch 7-31-2000, product index ANAEPG, with an ethylene-based solvent) in the salt feed slurry was maintained at 50 ppmv in the CSTRs and 100 ppmv in the Slurry Concentrating Tank by continuously pumping with syringe pumps. As in the previous tests, the concentration of  $^{137}\text{Cs}$  in the salt feed (average SRS recipe) was ~8 mCi/L, which was diluted by other feeds to ~6.5 mCi/L of slurry in the CSTRs.

### 2.2 FEED PREPARATIONS

Table 1 gives the masses of SRS average salt feed compounds needed to prepare the 140 L of salt feed for the test. Table 2 gives the masses of the chemical compounds needed to prepare 37 L of synthetic recycle wash water. This was added to CSTR 1 along with the other feeds to simulate the recycle of wash water from the TPB recovery and recycle process. Some actual TPB recycle wash water from Test 3 had been reserved for use at the beginning of Test 4. The quantity of synthetic wash water prepared assumed a 90-h run time at 6.86 mL/min after the recycle wash water from Test 3 was exhausted. The results of the washing of Test 3 slurry are shown in Table 3. Because of the low recovery of sodium tetraphenylborate (NaTPB) in the slurry wash, extra NaTPB was added directly to the recycle wash-water tank in the cell (to bring the TPB concentration up to 0.033 M) before the test was started. The synthetic recycle water contained enough NaTPB (0.03303 M) to provide 60% excess of the stoichiometric NaTPB requirement to CSTR 1. To provide a stoichiometric amount of NaTPB to CSTR 1 to precipitate the potassium, 6.52 L of the NaTPB feed was planned for Test 4 (0.68-mL/min feed rate). The concentrations of NaTPB and NaOH in this feed solution were 0.5545 and 0.96 M, respectively.

**Table 1. Average SRS salt feed preparation for Test 4**

Chemical species	Molarity (mol/L)	Compound	Molecular weight (g/mol)	Mass (g/L)	Mass <sup>a,b</sup> (g/206.4 L)
Cs <sup>+</sup> <sup>c</sup>	0.00016	CsCl	168.36	0.027	5.6
K <sup>+</sup>	0.0171	KNO <sub>3</sub>	101.1	1.729	356.9
OH <sup>-</sup>	3.611	NaOH	40.00	144.46	29,816.5
NO <sub>3</sub> <sup>-</sup>	1.388 <sup>d</sup>	NaNO <sub>3</sub>	84.99	117.97	24,349.0
NO <sub>2</sub> <sup>-</sup>	0.594	NaNO <sub>2</sub>	69.00	40.98	8,458.3
AlO <sub>2</sub> <sup>-</sup>	0.354	Al(NO <sub>3</sub> ) <sub>3</sub> · 9H <sub>2</sub> O	375.14	132.80	27,409.9
CO <sub>3</sub> <sup>2-</sup>	0.183	Na <sub>2</sub> CO <sub>3</sub> · H <sub>2</sub> O	124.01	22.69	4,683.2
SO <sub>4</sub> <sup>2-</sup>	0.171	Na <sub>2</sub> SO <sub>4</sub>	142.04	24.29	5,013.5
Cl <sup>-</sup>	0.0286	NaCl	58.4	1.671	344.8
F <sup>-</sup>	0.0366	NaF	41.99	1.537	317.3
HPO <sub>4</sub> <sup>2-</sup>	0.0114	Na <sub>2</sub> HPO <sub>4</sub> · 7H <sub>2</sub> O	268.09	3.056	630.7
C <sub>2</sub> O <sub>4</sub> <sup>2-</sup>	0.0091	Na <sub>2</sub> C <sub>2</sub> O <sub>4</sub>	134.00	1.219	250.6
SiO <sub>3</sub> <sup>2-</sup>	0.0046	Na <sub>2</sub> SiO <sub>3</sub> · 9H <sub>2</sub> O	284.2	1.307	269.8
MoO <sub>4</sub> <sup>2-</sup>	0.00023	Na <sub>2</sub> MoO <sub>4</sub> · 2H <sub>2</sub> O	241.95	0.056	11.5
H <sub>2</sub> O				796.0	164,290
Na <sup>+</sup>	6.42				
			Totals	1,289.8	266,207.6

<sup>a</sup>Assumes 160-h run time at 21.5 mL(salt feed)/min, or 1.29-L/h salt feed; also accounts for salt feed added to CSTRs and Slurry Concentrating Tank at beginning of test.

<sup>b</sup>Calculated density for preparation = (180,567 g)/(140,000 mL) = 1.29 g/mL.

<sup>c</sup>5.53 g of CsCl in 70 mL water was mixed well with 1.83 Ci of <sup>137</sup>Cs tracer (mass of <sup>137</sup>Cs was 0.021 g, but total cesium mass in tracer was 0.085 g). The <sup>137</sup>Cs concentration in the salt feed was 8.9 mCi/L.

<sup>d</sup>The target molarity for NO<sub>3</sub><sup>-</sup> is 2.45 mol/L. The addition of Al(NO<sub>3</sub>)<sub>3</sub>·9H<sub>2</sub>O provides 1.062 mol NO<sub>3</sub><sup>-</sup>/L [2.45 – 1.068 = 1.388 mol (NaNO<sub>3</sub>)/L extra needed].

**Table 2. Simulated recycle solution preparation for Test 4**

Chemical species	Molarity (mol/L)	Compound	Molecular weight (g/mol)	Mass (g/L)	Mass <sup>a,b</sup> (g/37 L)
TPB <sup>-</sup>	0.03303	NaTPB	342.23	11.30	418.10
OH <sup>-</sup>	0.6584	NaOH	40.00	26.346	974.80
NO <sub>3</sub> <sup>-</sup>	0.2531	NaNO <sub>3</sub>	84.99	21.511	795.91
NO <sub>2</sub> <sup>-</sup>	0.1083	NaNO <sub>2</sub>	69.00	7.473	276.50
AlO <sub>2</sub> <sup>-</sup>	0.0645	Al(NO <sub>3</sub> ) <sub>3</sub> · 9H <sub>2</sub> O	375.14	24.196	895.25
CO <sub>3</sub> <sup>2-</sup>	0.0334	Na <sub>2</sub> CO <sub>3</sub> · H <sub>2</sub> O	124.01	4.142	153.25
SO <sub>4</sub> <sup>2-</sup>	0.0312	Na <sub>2</sub> SO <sub>4</sub>	142.04	4.432	163.98
Cl <sup>-</sup>	0.0052	NaCl	58.4	0.304	11.24
F <sup>-</sup>	0.0067	NaF	41.99	0.281	10.40
HPO <sub>4</sub> <sup>2-</sup>	0.0015	Na <sub>2</sub> HPO <sub>4</sub> · 7H <sub>2</sub> O	268.09	0.402	14.87
C <sub>2</sub> O <sub>4</sub> <sup>2-</sup>	0.0017	Na <sub>2</sub> C <sub>2</sub> O <sub>4</sub>	134.00	0.228	8.44
SiO <sub>3</sub> <sup>2-</sup>	0.0008	Na <sub>2</sub> SiO <sub>3</sub> · 9H <sub>2</sub> O	284.2	0.227	8.40
MoO <sub>4</sub> <sup>2-</sup>	0.00004	Na <sub>2</sub> MoO <sub>4</sub> · 2H <sub>2</sub> O	241.95	0.0097	0.36
Na <sup>+</sup>	1.20				
H <sub>2</sub> O				950	35,150
Totals				1,051.12	38,881.50

<sup>a</sup>Assumes 90-h run time at 6.86 mL/min (or 0.412 L/h).

<sup>b</sup>Calculated density for preparation = (38,881.5 g)/(37,000 mL) = 1.05 g/mL.

**Table 3. Results for Test 3 slurry washing at 15 mL/min**

Washing volume (L)	Washing time (h)	Recycle wash water sample ID	RMAL NO <sub>2</sub> <sup>-</sup> (mg/L)	RMAL NaTPB (mg/L)
0	0	00T3W-1	20,800	267
3.6	4	04T3W-2	14,400	510
7.2	8	08T3W-3	9,130	850
10.8	12	12T3W-4	6,470	1,300
14.4	16	16T3W-5	3,630	1,140
18	20	20T3W-6	1,800	1,120
21.6	24	24T3W-7	1,440	948
25.2	28	28T3W-8	971	1,440
28.8	32	32T3W-9	721	1,260
Overall tank sample		36T3W-10	5,740	1,120

A total of 1237.2 g NaTPB [(0.5545 mol/L)×(6.52 L)×(342.23 g/mol)] and 25.05 g NaOH [(0.096 mol/L)×(6.52 L)×(40 g/mol)] were needed for the preparation. The concentration of MST in the slurry solution in CSTR 1 was set at about 0.5 g/L. The procedure for the MST stock solution preparation is given in CERS/SR/TPB/005.<sup>4</sup>

Test 4 was scheduled to run about 130 h, which provided for about 18 CSTR volume turnovers. The two CSTRs were left inventoried with process fluids from Test 3. Foam suppressant (IITB52) was added at a rate to provide concentrations of 50 ppm in the CSTRs and 100 ppm in the Slurry Concentrating Tank during Test 4. The various feed materials and feed rates for Test 4 are shown below:

Feed materials	Feed rates		
	Planned		Actual
	(mL/min)	(mL/h)	(mL/min)
Salt feed	21.5	1,209	21.64
Recycle wash water with excess NaTPB	6.86	411.6	6.95
MST + water (including MST, mercury, and palladium catalysts)	2.14	128.4	2.65
NaTPB (including 1PB and benzene)	0.68	40.8	0.73
IITB52 feed to CSTR 1 and CSTR 2	0.00156	0.0935	0.00156
IITB52 feed to Slurry Concentrating Tank	0.00312	0.1871	0.00312
Total feed rate	31.19		31.97

Catalyst feed	Feed rates		
	Planned		Actual
	(mg/L)	(mg/h)	(mg/h)
Mercury	85	159	136
Palladium	7.8	14.6	12.3/8.9 <sup>a</sup>
MST	500	935.7	1123
1PB	500	935.7	1005
Benzene	720	1,347.4	

<sup>a</sup>First half of test/second half of test.

### 3. MEASUREMENTS AND SAMPLING

#### 3.1 SAMPLING PLANS AND METHODS

During the tests, samples were taken for analysis according to the sampling plan outlined in CERS/SR/TPB/006, *Sampling Plan and Procedure (CSTR Test 4)*.<sup>5</sup> After preparation, the salt feed solution was analyzed for potassium, <sup>137</sup>Cs, and <sup>85</sup>Sr. During Test 4, effluent samples were taken for each CSTR (sample ports 1 and 2) and a filtrate sample was taken from the Slurry Concentrating Tank filtration system (sample port 4) every 4 h (a total of 95 samples of 55–65 mL each, collected over a 2-min duration). Each sample was analyzed using an on-site gamma counter for <sup>137</sup>Cs and <sup>85</sup>Sr. Every other sample (8-h time increment) was analyzed by the ORNL Radioactive Materials Analysis Laboratory (RMAL) for <sup>137</sup>Cs, <sup>85</sup>Sr, <sup>235</sup>U, potassium, boron, NaTPB, and TPB decomposition products. Other samples were archived in case additional data were needed.

Samples from ports 1 and 2 were prepared for analysis by filtering the whole sample with either a 0.45- $\mu$ m single filter or a 0.80/0.22- $\mu$ m double filter and placing filtrate in new sample bottles for transfer to the RMAL. (A double filter worked best for some of the samples because of the prefilter design of the syringe filter.) Reverse-phase high-performance liquid chromatography (HPLC) was used for NaTPB ( $\pm 20\%$ ), triphenylborane (3PB, ( $\pm 20\%$ )), diphenylborinic acid (2PB, ( $\pm 20\%$ )), phenylboronic acid (1PB, ( $\pm 20\%$ )), and phenol ( $\pm 20\%$ ). Gamma counting was used for cesium [<sup>137</sup>Cs ( $\pm 10\%$ )] and strontium [<sup>85</sup>Sr ( $\pm 10\%$ )]. Inductively coupled plasma-atomic emission spectroscopy (ICP-AES) was used for potassium ( $\pm 10\%$ ) and boron ( $\pm 10\%$ ) analyses. The potassium and boron samples were digested in nitric acid in a microwave oven prior to analyses. The filtrate samples from the Slurry Concentrating Tank (sample port 4) were sent as-collected to RMAL for analysis after a small sample (1–3 mL) was placed in a counting tube for on-site <sup>137</sup>Cs and <sup>85</sup>Sr counting, which had an error margin of  $\pm 10\%$  and minimum detectable levels of 3E-5  $\mu$ Ci/g for both <sup>85</sup>Sr and <sup>137</sup>Cs. All samples destined for the RMAL were placed in a refrigerator after preparation, until they could be transported. A chain-of-custody procedure in accordance with the Sample Management Office was followed for all samples sent to RMAL.

The <sup>85</sup>Sr in the Test 4 feed resulted from the small amount of simulant intentionally left in the feed tank from Test 3. Almost one-half of the <sup>85</sup>Sr had decayed since Test 3, and the Test 3 simulant left in the feed tanks was diluted with new feed by a ratio of about 8 to 1.

#### 3.2 MEASURING AND TEST EQUIPMENT

Electronic balances, which are calibrated annually by ORNL Plant and Equipment Division (P&E) personnel, were used for solid and liquid mass measurements in preparing the simulants. Class A volumetric flasks and graduated cylinders were used for simulant preparation. To monitor levels of <sup>137</sup>Cs and <sup>85</sup>Sr during the CSTR tests, a gamma-counting system consisting of an ORTEC model GMX-45220-P-S intrinsic germanium detector, an ORTEC model 672 counting system amplifier, a Canberra Accuspec-A MCA card, and Canberra Genie-2000 spectroscopy software were used.

Instrumentation used in the CSTR system was calibrated before startup. Calibration records are maintained in registered logbooks and in a controlled project file. A revised list of instruments and corresponding calibration schedules was generated after the modifications to the CSTR system had been reviewed and approved by SRS and Oak Ridge investigators (CERS/SRS/TPB/009, 8/04/2000).

## 4. RESULTS OF TEST 4

### 4.1 RUN SUMMARY

Test 4 was successfully conducted October 9–15, 2000. The actual run time (time in which the feeds were being pumped to CSTR 1) was 125.5 h. About 18 h after starting the test, the run was put on hold for 5 h to replace the CSTR 1 sample valve, which had plugged and allowed the level in CSTR 1 to rise almost to the overflow level. Two batches of concentrated slurry (about 10 wt % insoluble TPB and MST reaction products) were produced and then washed to evaluate the recovery of TPB after the test was completed. The temperature of the slurries in the system vessels was maintained at 25°C for the first 76 h of operation, producing the first batch of 10 wt % concentrated slurry. The temperature was then ramped to 45°C over the next 13 h. During this period, the temperature in each vessel was increased in 5°C steps and held for about 3 h at each step. The temperature was maintained at 45°C for 43 h to produce the second batch of 10 wt % concentrated slurry. Benzene was not added to the system during the testing at elevated temperature.

The test plan (CERS/SR/TPB/010, Rev. 0)<sup>6</sup> called for initial operation at 25°C for the production of the first batch of concentrated TPB/MST slurry. The two CSTRs were full of slurry that was left in these vessels at the completion of Test 3, with the intent of maintaining the required cesium DF from the beginning to the end of the test. To prevent settling of the solids from Test 3, the agitators in both CSTRs were operated at 500–600 rpm for the 2-month period between Test 3 and Test 4 startup. The vessel agitators were then operated at 650–850 rpm throughout Test 4. Although the test plan called for operation at up to 1200 rpm, the speeds were not increased after the first valve plug incident in an attempt to prevent any solids from plugging the slurry transport lines connecting the CSTRs and the Slurry Concentrating Tank. Antifoam (IITB52) was pumped to each of the vessels during the test. The test plan called for the MST/water feed supply containing the catalyst to provide concentrations of about 7.8 mg of palladium and 85 mg of mercury per liter of slurry in CSTR 1. A solution of 1PB was added to the NaTPB feed stream and provided about 500 mg of 1PB per liter of slurry in CSTR 1. Benzene was added to the TPB feed stream (via a syringe pump that was connected to the feed line) at a rate that provided a benzene concentration in the CSTR 1 slurry of about 720 mg/L. The benzene addition began 23 h after the test started and was discontinued prior to increasing the process temperature to 45°C. Benzene production was monitored in each of the process vessels by on-line FTIR analyses of the nitrogen purge gas for the vessels.

The process control of the slurry concentration system for Test 4 was excellent. The cross-flow filter backpulsing system worked well and was very effective at 25 and 45°C. No backpulse was needed after an initial one near the end of the temperature ramping period. No foam problems occurred in either the CSTR or in the Slurry Concentrating Tank. The benzene analysis system generally worked well and allowed the benzene in the vessel vapor spaces to be continually monitored during all parts of the test. The slurry wash cross-flow filter also worked well, and no problems occurred during the slurry washes.

Overall, more than 161 L of radioactive simulated feed was treated. Two 6–7 L batches of 10 wt % concentrated slurry were collected and subsequently washed. More than 230 L of filtrate was collected from the Slurry Concentrating Tank cross-flow filter. Each batch of concentrate was washed for 24 h, and the separated wash water (~28 L) was collected and saved for use as startup recycle water in Test 5. The planned and actual flow summaries are shown in Table 4. Afterward, both cross-flow filter systems were chemically cleaned with 2% oxalic acid solution, rinsed well with water, and then rinsed and stored in 1 M NaOH.

The overall system operation and the behavior of the components of the system met test plan specifications and objectives. The most difficult operational problems in Test 4 occurred with the MST/catalyst/dilution water delivery system. One was maintaining the MST and catalyst particles in a uniform suspension in the feed tank. The mixture of MST, mercury, and palladium (5%) on gamma alumina proved significantly more dense and “sticky” (i.e., it agglomerated to form stringy particle masses that settled quickly) than the MST/water-only slurry had been in previous tests. The other problem occurred in the transfer tubing, which often plugged when solids settled at wide places in the line or in the check valve in the line inside the hot cell. Finally, a plug in the check valve could not be cleared and flow was lost for 3–4 h (between the hours 98 and 104 of the test). To continue MST/catalyst/dilution water flow to CSTR 1, a plumbing connection was made with the TPB line at a tee next to the cell wall so that the MST/catalyst/dilution water could be added with the TPB solution. With time, the check valve for this line even tended to plug; however, pumping continued until the end of the test without further interruptions to either the TPB or the MST/catalyst/water flow.

**Table 4. Process performance for Test 4**

	Volume (L)	
	Planned <sup>a</sup>	Actual <sup>b</sup>
Salt feed	161.9	167.6 (~4.6 L leaked out during last 10 h)
Recycle wash water	51.5	24.2 actual + 28.1 synthetic = 52.3 total
Na TPB	5.1	5.5
Water + MST	16.1	20.0 <sup>c</sup>
Filtrate from cross-flow filter	218.9	218
Slurry produced	13	13

<sup>a</sup>Planned test time was 132 h for both parts.

<sup>b</sup>Actual test time was 130.3 h (125.5 h actual feeding).

<sup>c</sup>The flow was off for 4 h (water + MST+ catalyst) when the feed line plugged in the check valve. Flow was rerouted through the TPB feed line.

For Test 5, modifications to the MST/catalyst/dilution water delivery system must be made to mitigate or prevent the reoccurrence of the Test 4 problems. The feed lines will be replaced, and the feed tank/agitator system must be redesigned. In addition, an effort will be made to determine chemical adjustments necessary to prevent agglomeration of the MST/catalyst mixture.

## 4.2 DETAILS OF SYSTEM CHEMISTRY

### 4.2.1 Cesium and Strontium Decontamination of Test 4 Feed

During Test 4, even at elevated temperatures, the <sup>137</sup>Cs DFs for the filtrates from the Slurry Concentrating Tank (as measured by ORNL RMAL) were >56,000, with some values as high as 650,000. The decontamination chemistry was excellent; these values exceeded the DF goal of 40,000 (99.998% removal). Furthermore, DF values for filtrate samples from CSTR 2 were also very high, ranging from 21,000 to 104,000, with most values >37,000. The same was the case for the CSTR 1 filtrate samples,

which had DFs ranging from 21,000 to 140,000. A high percentage of the samples had values >30,000. The  $^{137}\text{Cs}$  DF results are given in Figs. 7–12 and provide a comparison with the on-site counting results. It should be noted that most of the cesium decontamination appeared to be completed in the first CSTR, but this may be an artifact of the sampling. Because at least 15 min is required between the time the sample is removed from the CSTR and the time that the filtrate and slurry are separated, additional precipitation time for the cesium removal is available. The need for the second CSTR is primarily for the removal of strontium and uranium by MST. The use of the palladium catalyst at 45°C did not cause any loss of DF (in any vessel) even with the TPB actively decomposing.

The on-site process control analysis gave consistently higher values for cesium (or lower DFs) than did RMAL analysis. This is partially due to differences in counting geometry, etc. The  $^{137}\text{Cs}$  DFs for the Slurry Concentrating Tank filtrates ranged between 16,000 and 287,000, with most values >30,000. During the first part of the test (at 25°C), the CSTR 1 on-site and RMAL results are close. However, during the higher-temperature portion of the test, the difference between the two analyses is somewhat greater. Figure 11 shows little difference between the on-site and RMAL cesium removal results for CSTR 2. The on-site cesium DFs were in the 17,000 to 63,000 range, with most values >25,000. The sampling results for the Slurry Concentrating Tank filtrates are given in Fig. 12. These show a wider variation between the on-site and RMAL analyses. The differences were smaller when the slurry was being heated during the last 55 h of the test. The DF values for the Slurry Concentrating Tank filtrates are the most important. They show that the process, even with TPB decomposition occurring, was producing filtrate that greatly exceeded the SRS WAC requirements for saltstone.

The on-site cesium DFs were in the 17,000 to 63,000 range during the time between completion of the first batch of 10 wt % slurry and the end of the test during operation at 45°C (see Fig. 9). Several filtrate samples indicated (by on-site counting) reduced DFs in the range of 3000 to 10,000. Samples of those filtrates that were refiltered at room temperature with 0.2- $\mu\text{m}$  syringe filters gave DFs that were higher and in the acceptable range. This indicates that insoluble  $^{137}\text{Cs}$  penetrated the cross-flow filter membrane or was precipitated after the cross-flow filtration and cooldown of the samples.

No strontium was added to the salt feed for Test 4. However, residual  $^{85}\text{Sr}$  remained in the heel of the feed left from Test 3, which was designed to measure the strontium removal. The calculated strontium concentration in the Test 4 salt feed was 5.6  $\mu\text{g/L}$ , compared with 44.8  $\mu\text{g/L}$  in the Test 3 salt feed. The  $^{85}\text{Sr}$  data plotted in Fig. 13 show that the MST effectively removed strontium in Test 4; DFs in the range of 100 to 200 were maintained for the duration of the test (above the DF level of approximately 26 that is anticipated to be required).

Although the palladium catalyst mixture was continuously added to cause TPB decomposition during Test 4, the cesium DFs for CSTR 1 remained in the 21,000 to 140,000 range, as shown in Fig. 10. Furthermore, the cesium DF was not lost during the time between Test 3 and Test 4, as shown by the “time-zero” samples. In addition, cesium removal was not affected by increasing the operating temperature to 45°C and operating at that temperature during the last 42 h of the test.



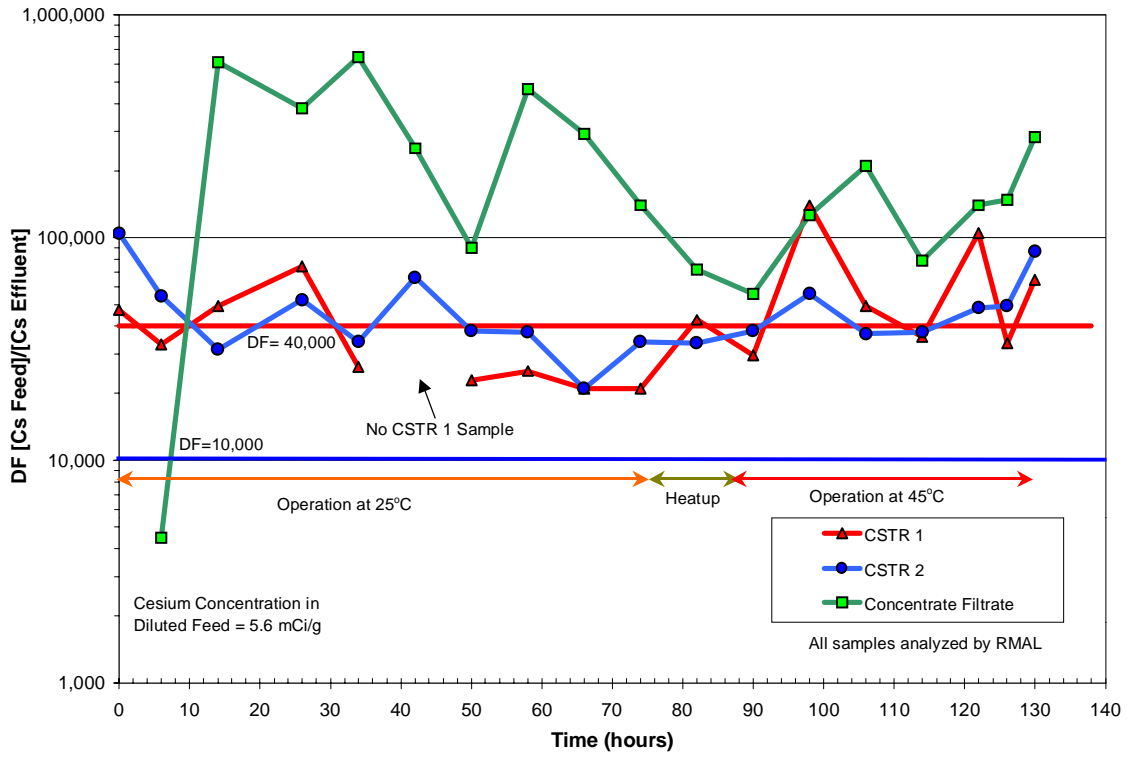


Fig. 7. Test 4 cesium DFs.

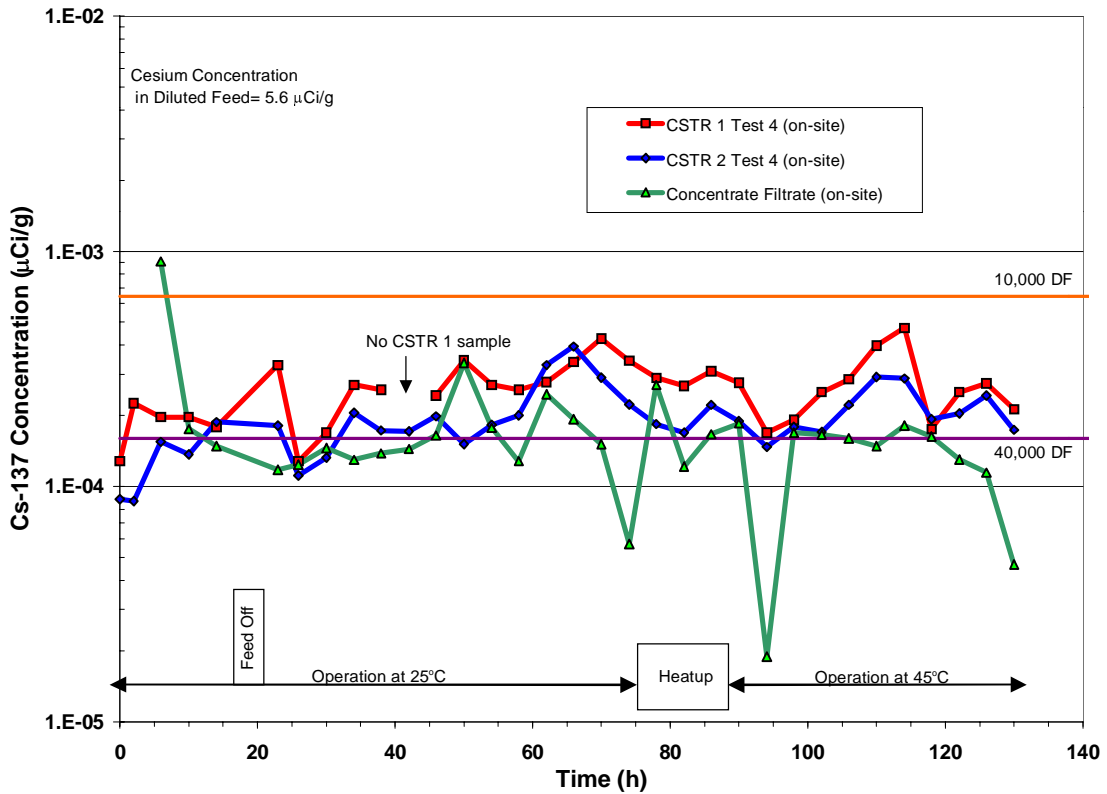


Fig. 8. Cesium concentrations in Test 4.

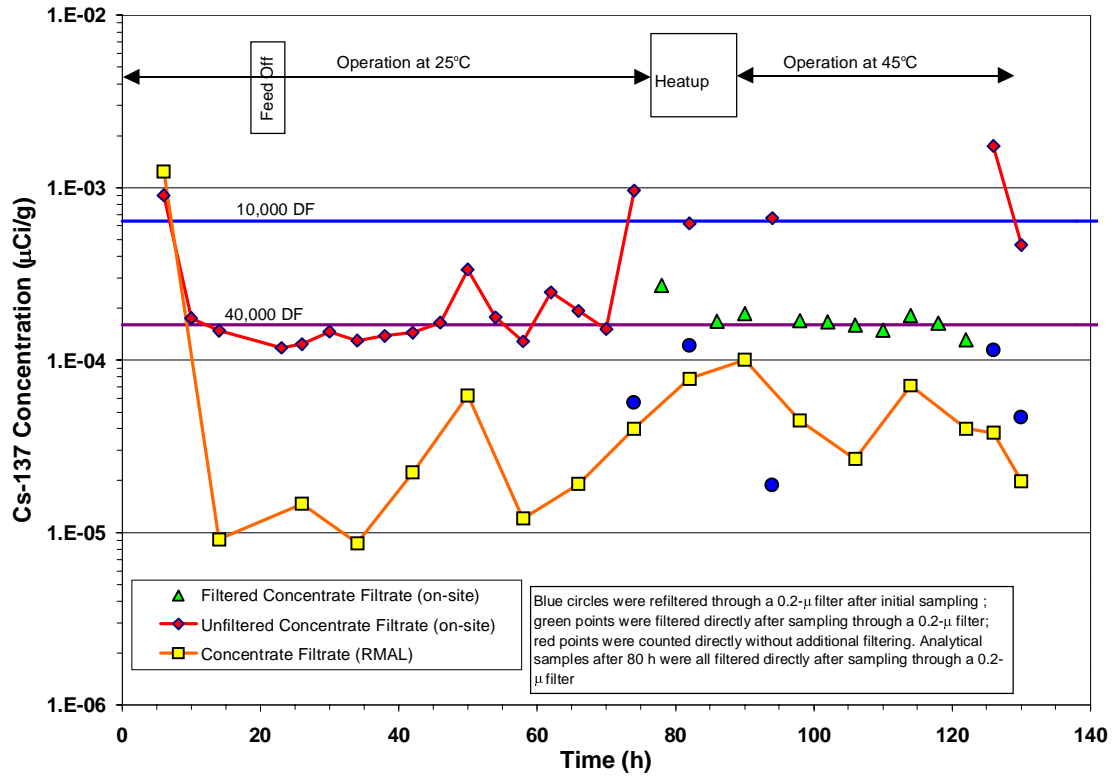


Fig. 9. Cesium concentrations in Test 4 concentrate filtrate showing filtration differences.

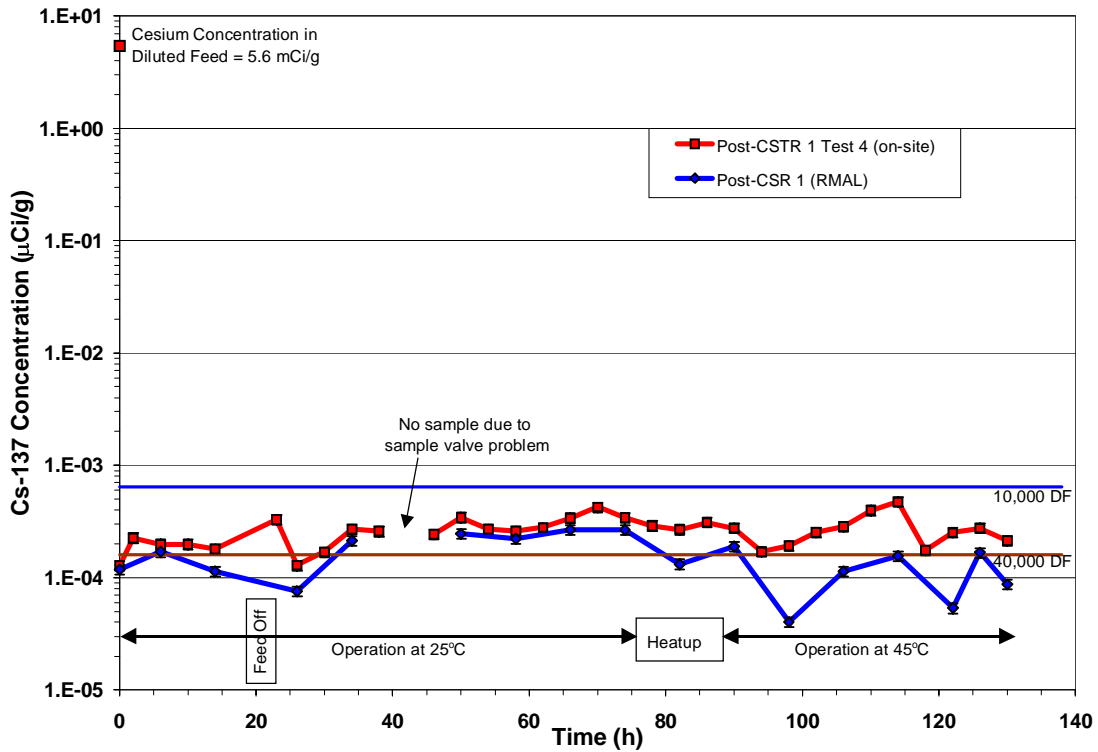


Fig. 10. Cesium concentrations in Test 4, after CSTR 1.

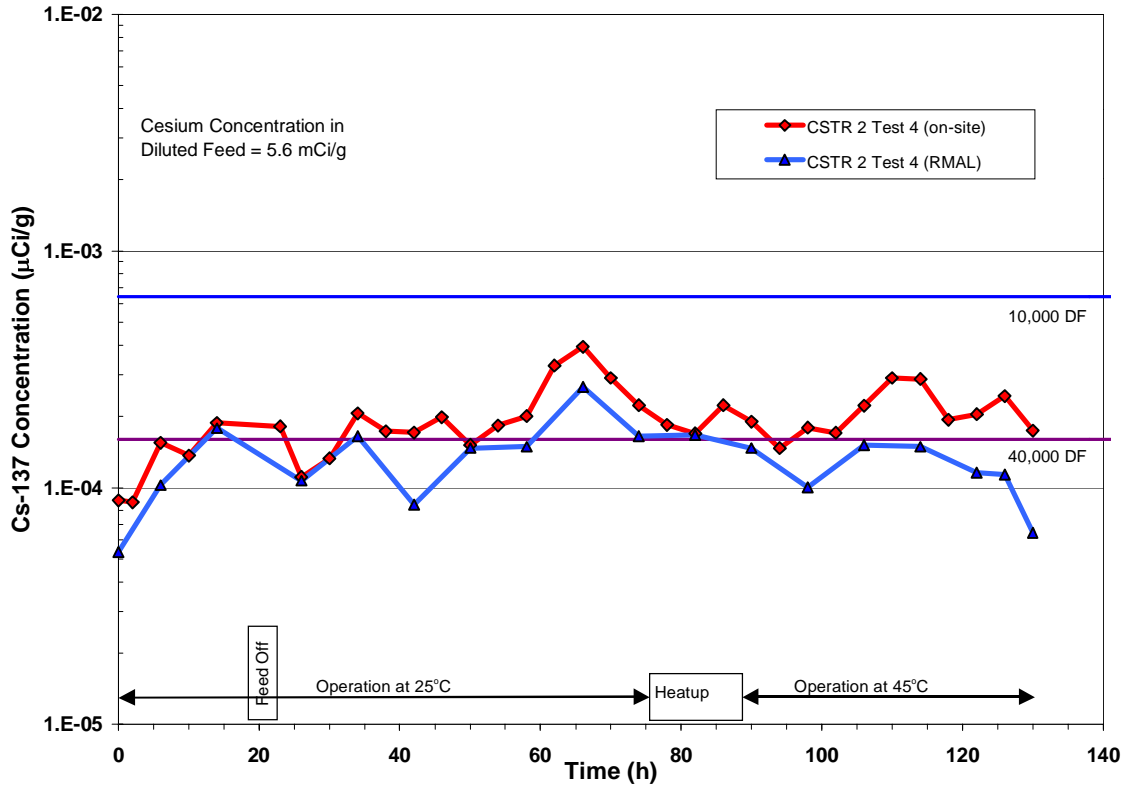


Fig. 11. Cesium concentrations in Test 4, after CSTR 2.

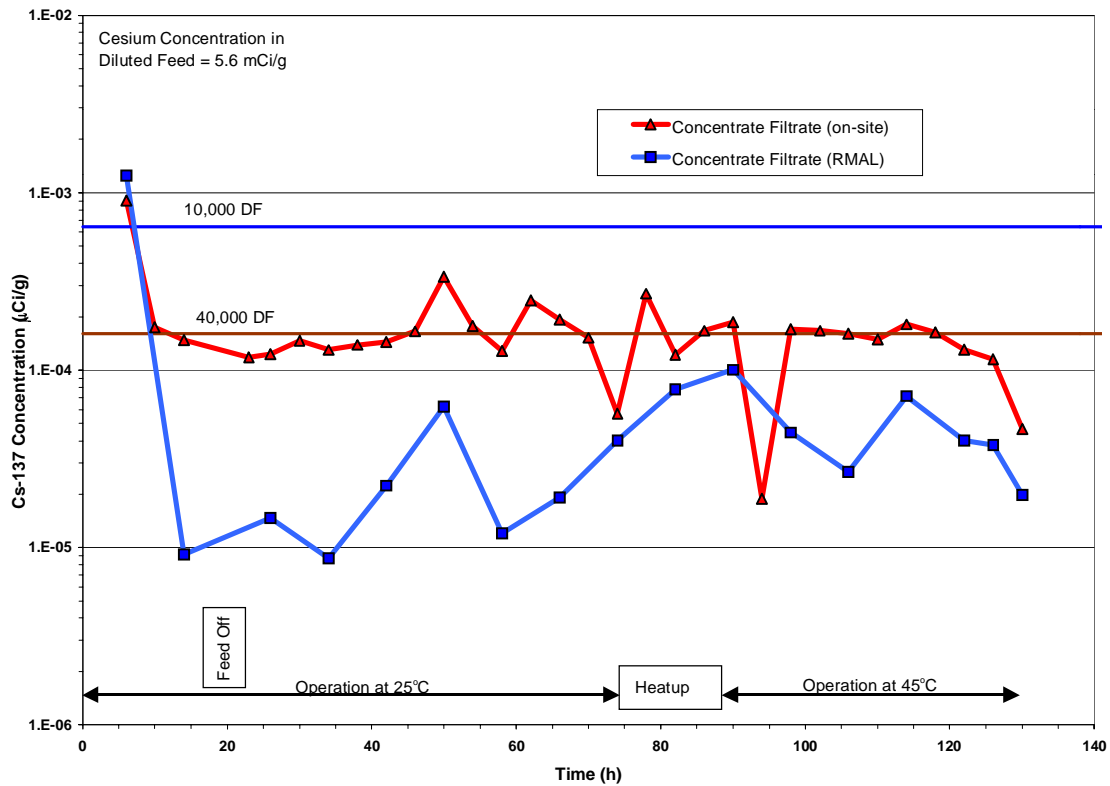


Fig. 12. Cesium concentrations in Test 4, after Slurry Concentrating Tank.

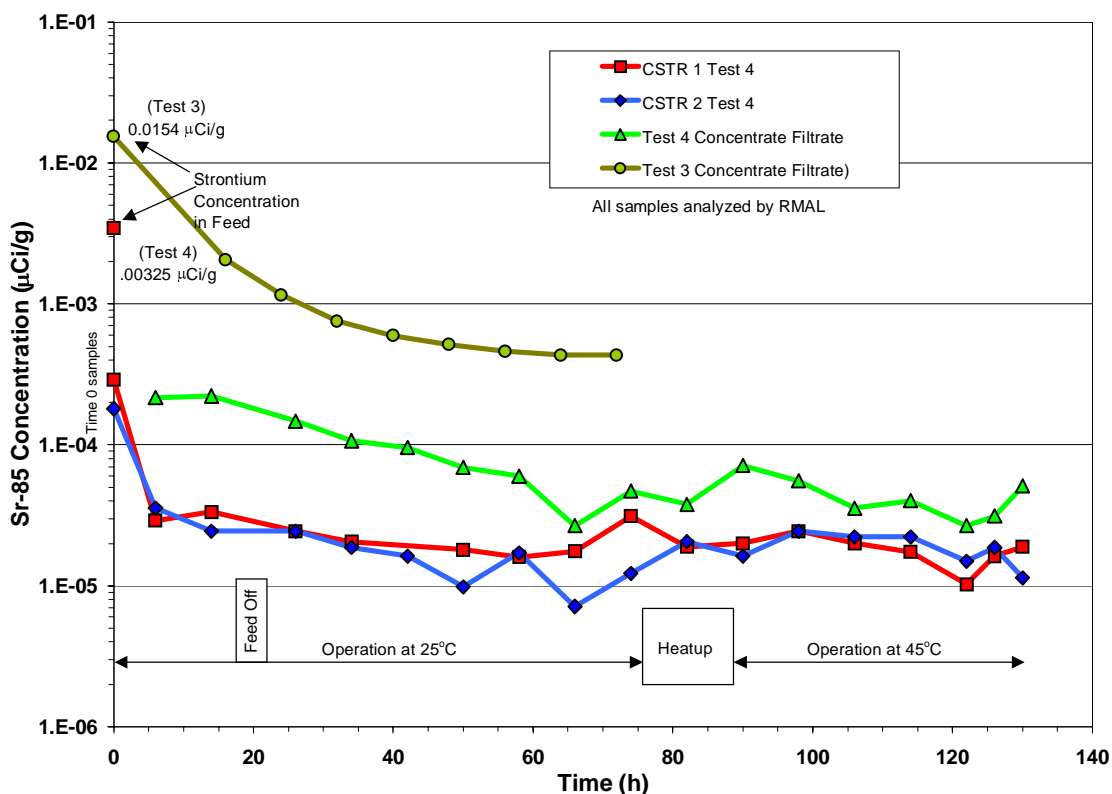


Fig. 13. Analytical concentrations of  $^{85}\text{Sr}$  in Test 4 process vessels.

#### 4.2.2 HPLC Analyses for TPB Decomposition Products

Figures 14–16 show the results of the HPLC and ICP analyses of samples of filtrates from the two CSTR filtered overflows and the concentrate filtrate. The graphs do not show any results for 2PB, because of difficulties in resolving its peak in the presence of the large amount of 1PB added to the system with the feed. Each figure also shows the benzene concentration of the vessel off-gas, as analyzed by the on-line FTIR analyzers. In Fig. 14, CSTR 1 analyses show the TPB decomposition rates calculated from the benzene in the off-gas during the addition of benzene for the 23- to 68-h time frame and from heating the slurry at 45°C in the last 47 h of the test (with no benzene addition). Benzene was pumped via the NaTPB feed line to CSTR 1 at a rate of 22.46 mg/min, which was calculated to yield a concentration of benzene in the slurry of 720 mg/L. No benzene was detected in the CSTR 1 off-gas until the benzene addition began. The benzene concentration in the off-gas increased during the next 10 h to about 400 ppm and remained at  $398 \pm 35$  ppm until the addition of benzene was stopped at 68 h. After this point, the off-gas concentration decreased to about 90 ppm during the next 18 h. The slurry was heated to 45°C during the last 12 h of this time. Even after the process temperature reached 45°C, the benzene concentration continued to decline to about 60 ppm over the next 12 h. It was found, however, that the catalyst/MST feed line had become plugged for about 3 h during the 95- to 98-h time period. At 101 h into the test, the catalyst/MST line was attached to the NaTPB supply line to continue providing MST and catalyst to the CSTR. To overcome the loss of MST and catalyst to the CSTR, the catalyst/MST flow was doubled for about 4 h. At this point in the test, the concentration of benzene, which could be generated only by TPB decomposition, dramatically increased to about 670 ppm at 118 h and remained at  $660 \pm 48$  ppm until the end of the test. This increase was likely due to the combined effect of temperature increase and increased catalyst feed rate.

The concentrations of TPB and TPB decomposition products (also shown in Figs. 17–19) followed reasonable trends in CSTR 1. Phenylboronic acid (1PB), which was pumped (via the NaTPB feed) to CSTR 1 at a rate of 15.6 mg/min, provided the targeted concentration of about 500 mg/L. During the first 14 h of the test, the concentration of 1PB in the CSTR 1 slurry filtrate, as seen in Fig. 14, increased into the range of 400 to 500 mg/L. The concentration remained in that range with the exception of two short-term dips that may be related to the analytical determinations and possible interferences from other components in the sample matrix. Through the first 78 h at 25°C, the concentration of 3PB in CSTR 1 averaged  $28 \pm 8$  mg/L (equates to 2.3% decomposition of the excess NaTPB). After 78 h, the concentration began to increase and continued to climb until the end of the test, reaching 159 mg/L (equates to 4.2% decomposition of the total TPB, or about 10% of the excess TPB). In the decomposition of TPB, 1 mol of benzene is formed for each mole of 3PB formed. The highest 3PB molarity obtained in CSTR 1 was 0.61 mM, which would also have been the concentration of benzene had it remained in liquid form in the slurry. The concentration of soluble TPB averaged  $57.8 \pm 24$  mg/L for the test. Its concentration was highest while the slurry was being heated, averaging  $77 \pm 17$  mg/L. Continuing to maintain TPB in solution guarantees that TPB is available to precipitate any cesium that is present. These data indicate the expected rate of TPB decomposition and the expected increase in decomposition rate when the temperature was increased.

Analyses of the filtered samples from CSTR 2 show somewhat similar results (Fig. 15). The 3PB concentrations were very low during the first 34 h of the test; all but two of the filtrate samples had values below the detection limit of 15.7 mg/L. From 34 until about 82 h into the test, the average 3PB concentration was  $27 \pm 13$  mg/L. This test period also included much of the slurry heat up ramp. Subsequently, the 3PB concentration increased steadily to about 140 ppm while the slurry was being heated at 45°C. Because the 3PB filtrate concentration in CSTR 2 was approximately equal to that in CSTR 1 and because the CSTR 2 slurry represents overflow from CSTR 1, there appears to be little additional decomposition of the TPB in CSTR 2 unless the decomposition rate of 3PB to 2PB equals the TPB decomposition rate to 3PB (true at steady state). The soluble TPB concentration averaged  $39 \pm 16$  mg/L. Benzene did not appear in the off-gas of CSTR 2 until after 42 h of testing. Its concentration steadily increased over the next 28 h to 200 ppm, about the time when the slurry heatup began. During the slurry heatup period, the benzene concentration decreased to about 90 ppm. Once at 45°C and for the remainder of the test, the benzene concentration steadily increased to about 350 ppm. These data show clear evidence of the expected trends of TPB decomposition.

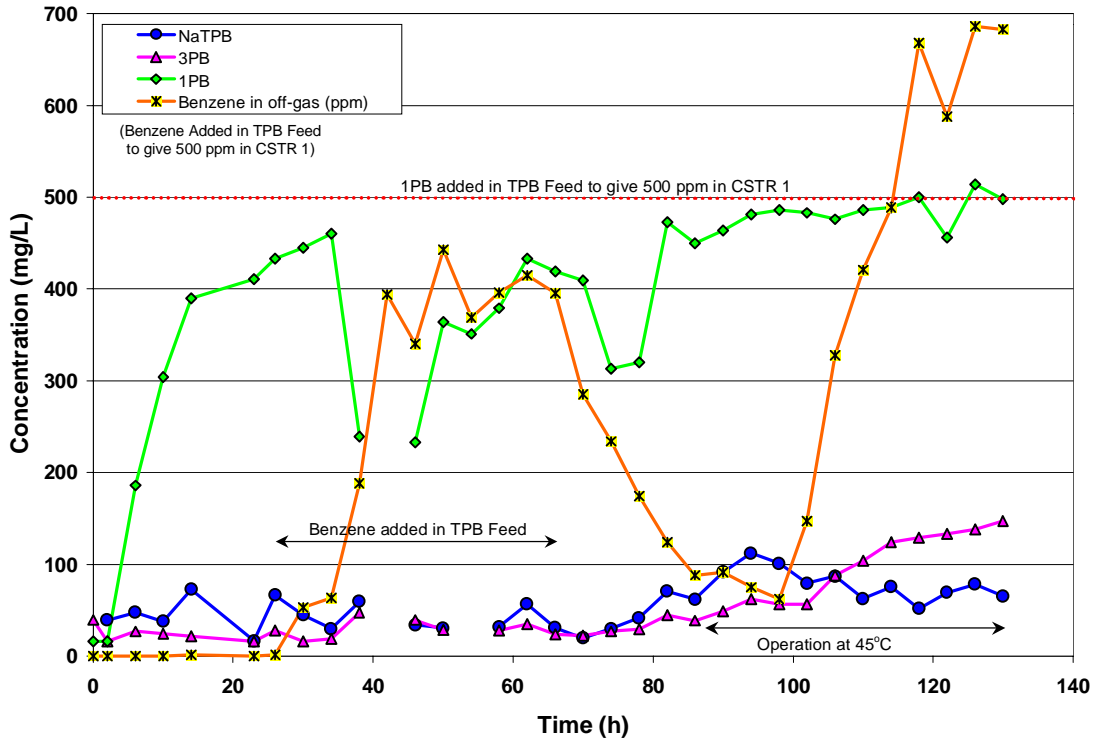


Fig. 14. CSTR 1 HPLC results for Test 4.

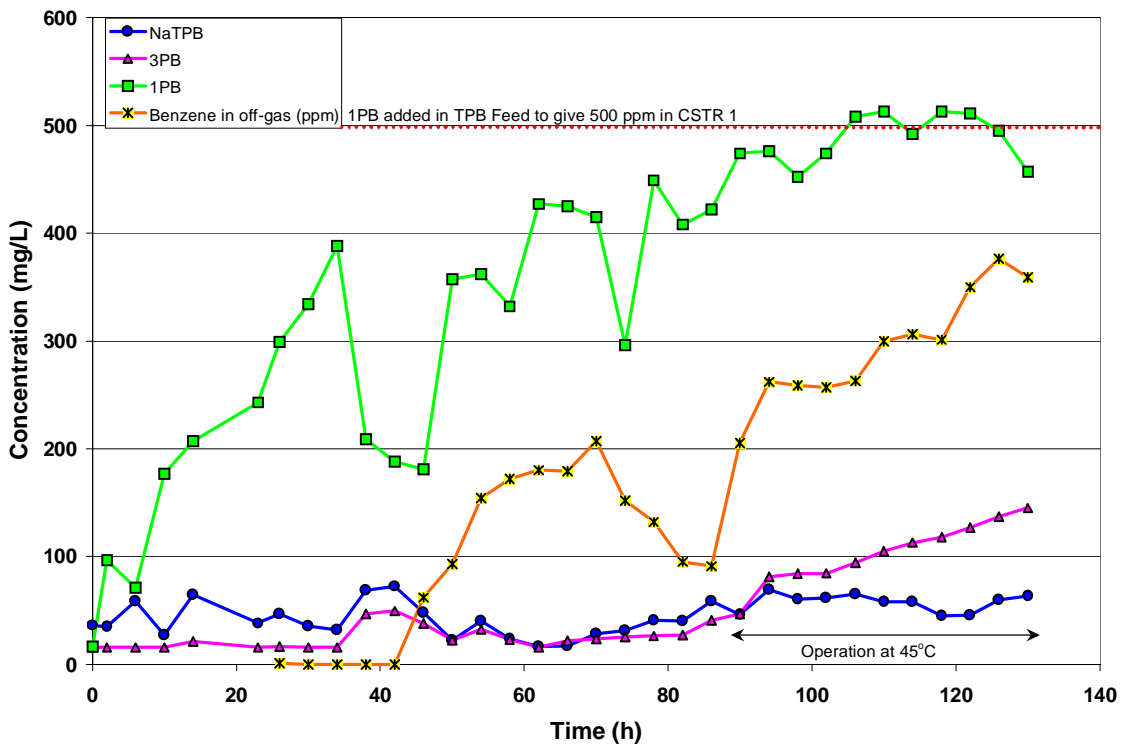


Fig. 15. CSTR 2 HPLC results for Test 4.

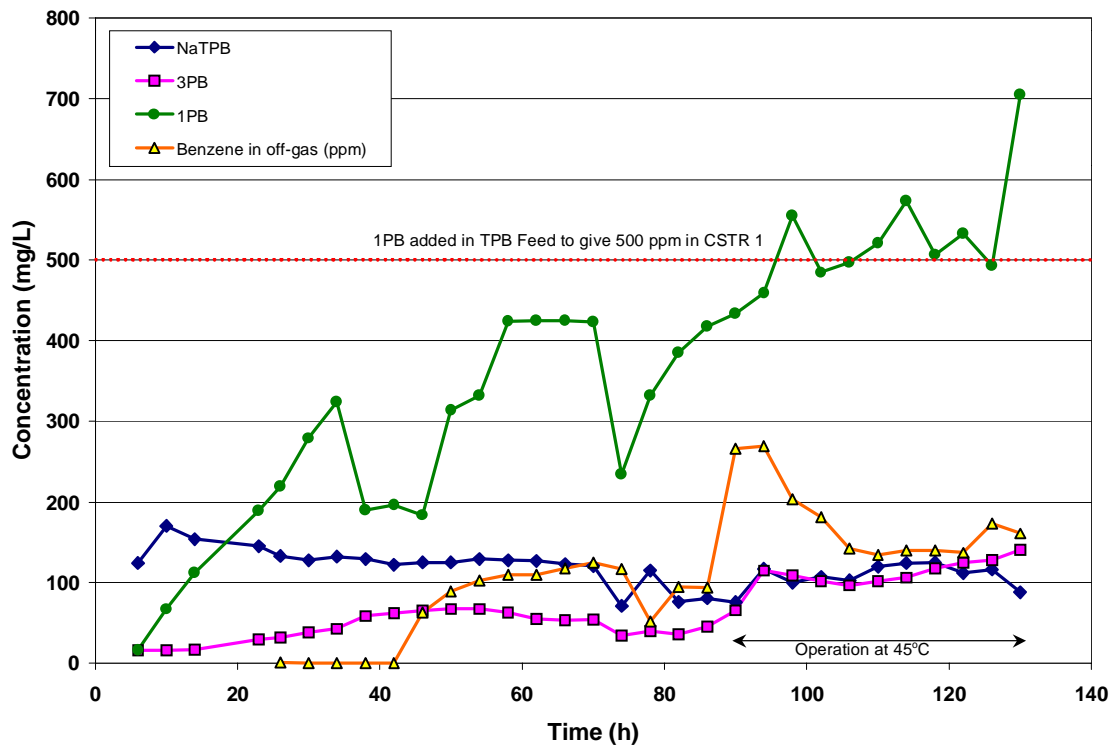


Fig. 16. Slurry Concentrating Tank HPLC results for Test 4.

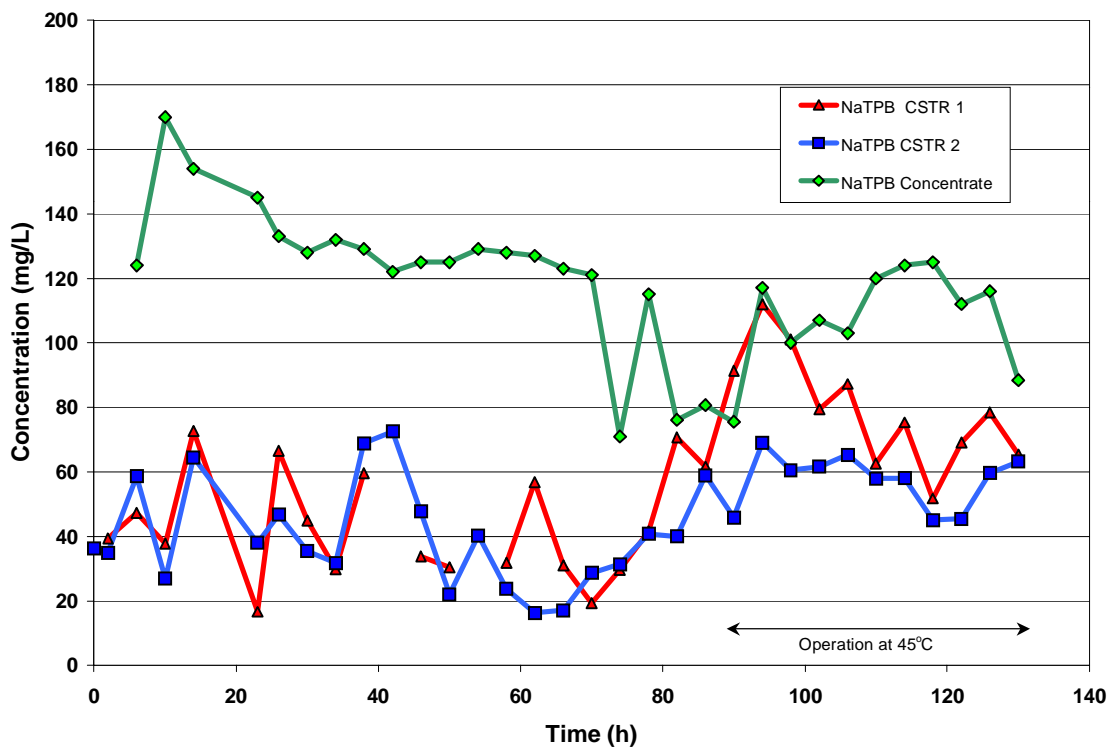


Fig. 17. NaTPB analysis results for Test 4.

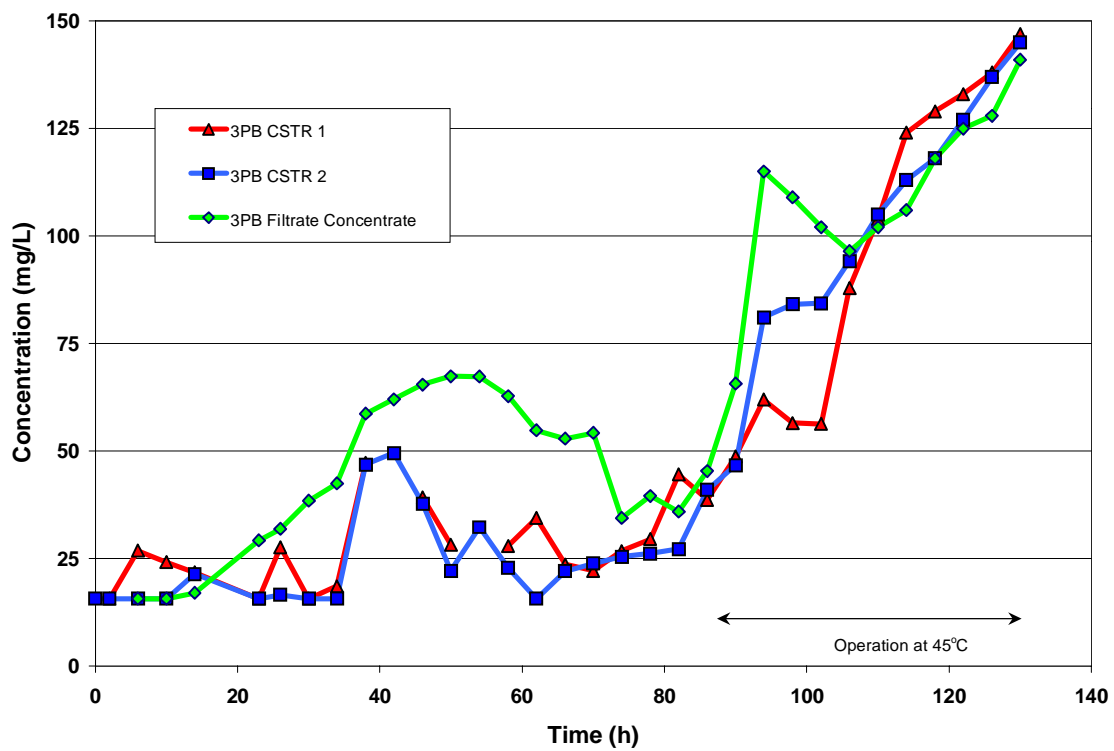


Fig. 18. 3PB analysis results for Test 4.

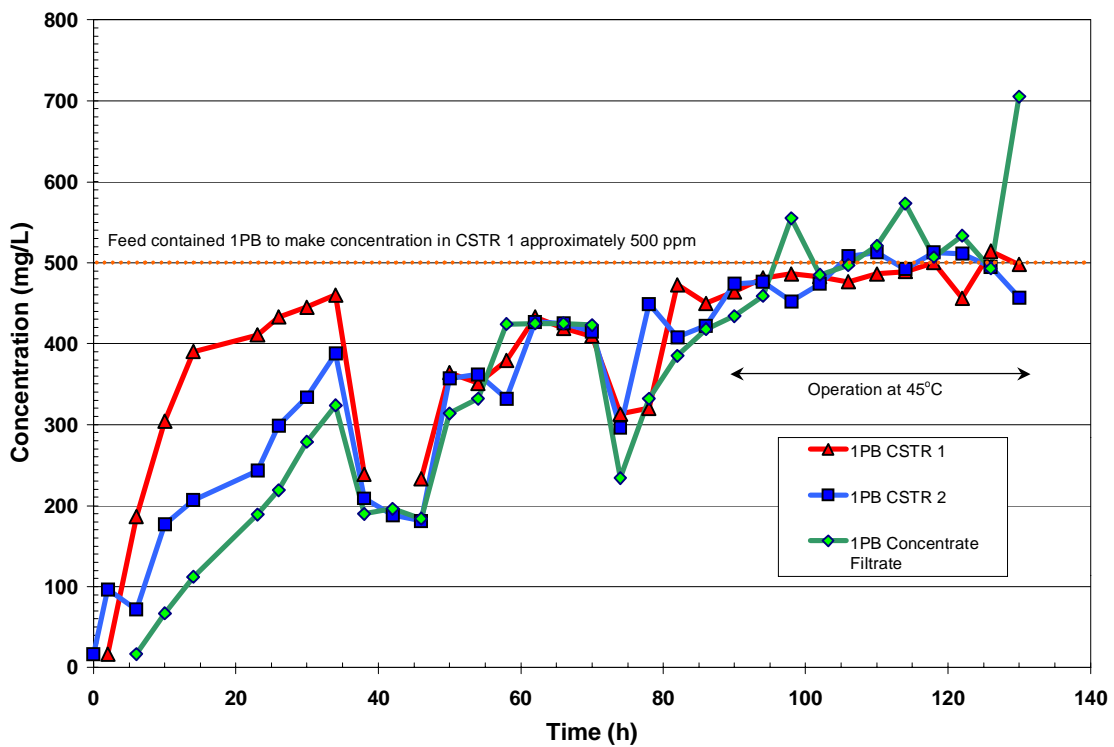


Fig. 19. 1PB analysis results for Test 4.



The HPLC results for the Slurry Concentrating Tank filtrate are provided in Fig. 16. During the first half of the test, while the TPB solids were being concentrated to 10 wt %, the concentration of soluble TPB averaged  $132 \pm 16$  mg/L. After these solids were transferred to the Wash System, the concentration decreased to as low as 71 mg/L. While the second batch of slurry was being concentrated, the TPB concentration increased to an average of  $114 \pm 9$  mg/L for the last 40 h of the test while operating at 45°C. The 3PB levels were below the detection limit during the first 10 h of the test but averaged  $50 \pm 16$  mg/L until the concentrated solids were pumped to the Wash System at 64 h into the test. During concentration of the second batch of slurry at 45°C, the 3PB concentration increased from a low of 20 mg/L to a high of 140 mg/L at the end of the test. The 2PB concentrations were below the detection limits. The concentration of 1PB increased to about 300–350 mg/L during the first 40 h of operation, decreased after slurry transfer, and then increased to a range of 500 to 600 mg/L during the last 32 h of the test at 45°C. No benzene was detected in the Slurry Concentrating Tank off-gas during the first 42 h of the test. Its concentration then increased to 100 ppm during the next 14 h and remained between 100 and 120 ppm until the temperature of the slurry reached 45°C at the 86th hour of the test. The benzene concentration increased rapidly after 86 h to 270 ppm for about 8 h, then slowly decreased to about 140 ppm at the 106th hour and remained between 140 and 170 ppm for the remainder of the test. Once again, the trends of TPB and TPB decomposition products indicate the expected rate of decomposition and the increasing trend in the decomposition rate as temperature is increased.

#### 4.2.3 ICP-AES Analyses for Potassium and Boron

Figures 20–23 show the ICP-AES results for potassium and boron for the filtered samples from CSTR 1, CSTR 2, and the filtrate from the slurry concentration filter. The potassium for all three vessels started out at lower levels (100–120 mg/L) than those measured at the end of Test 3 (about 200 mg/L). This indicates continued precipitation of potassium through reaction with any excess TPB that may have been present. Throughout the course of the test, potassium concentrations in all three vessels gradually decreased to approximately 50–60 mg/L by the end. Figures 20–23 also show the boron concentrations and the calculated boron concentrations (equates to the soluble TPB, 3PB, and 1PB and is indicated as PBs in the figures). As shown in Fig. 20, the boron concentration in CSTR 1 increased during the first 30 h of operation, was fairly constant at about 40 mg/L during the next 40 h, and then slowly increased to about 53 mg/L while the temperature was increased to and maintained at 45°C. The figures illustrate almost no difference between the concentrations of boron measured and those calculated based on the amounts of soluble TPB, 3PB and 1PB present in any of the vessels. The concentration of boron in Test 4 should be about 5 ppm based on the 1PB added with the feed and an additional 2–3 ppm for the 3PB created by decomposition. Since no copper was present to rapidly catalyze the conversion of 1PB to benzene, very little boric acid would be expected.

#### 4.2.4 Benzene Production Rate as a Measure of TPB Decomposition

An assessment of the TPB decomposition rate was made based on the concentrations of 3PB and 2PB in the filtrates from the CSTRs and the Slurry Concentrating Tank and from the benzene concentrations in the off-gas of these tanks. In this assessment, the decomposition of TPB is indicated by the production rate of benzene. During decomposition, each phenyl group that is removed from a TPB molecule can potentially produce a benzene molecule. If a TPB molecule is completely decomposed, four molecules of benzene could be produced. Calculated decomposition rates are illustrated in Figs. 24–28 as benzene generation rates (expressed as grams of benzene per liter of reactor volume per hour). Figure 24 shows the CSTR 1 results and gives the benzene rate based on the concentration of benzene in the vessel off-gas. The target rate for decomposition of TPB for the test was 10 mg benzene/(L·h) based on the complete decomposition of TPB. Because of the lack of catalyst for 1PB decomposition and the difficulty in analyzing for 2PB, decomposition rates are based primarily on the

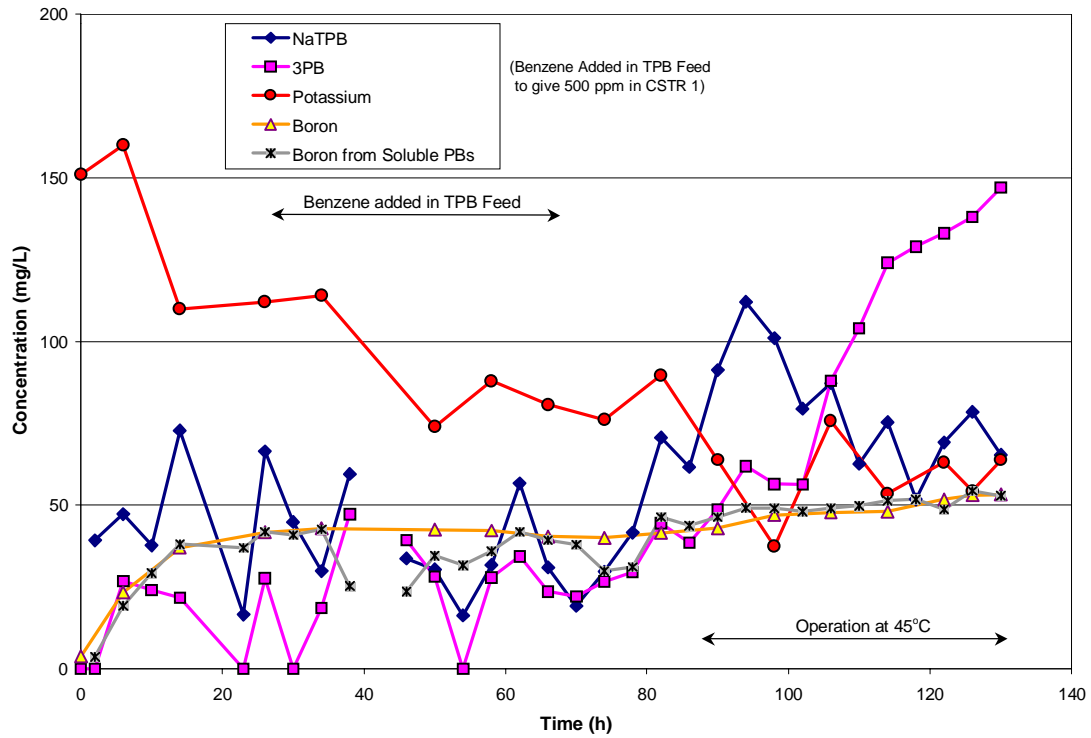


Fig. 20. CSTR 1 boron and potassium analysis results for Test 4.

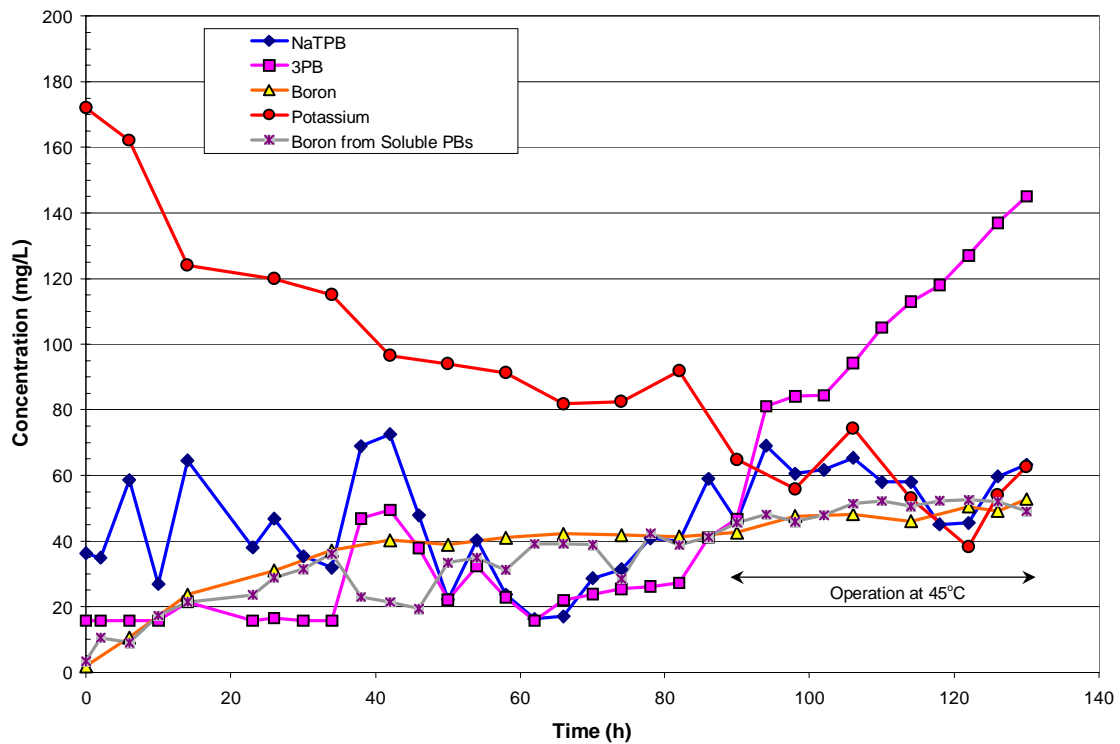


Fig. 21. CSTR 2 boron and potassium analysis results for Test 4.

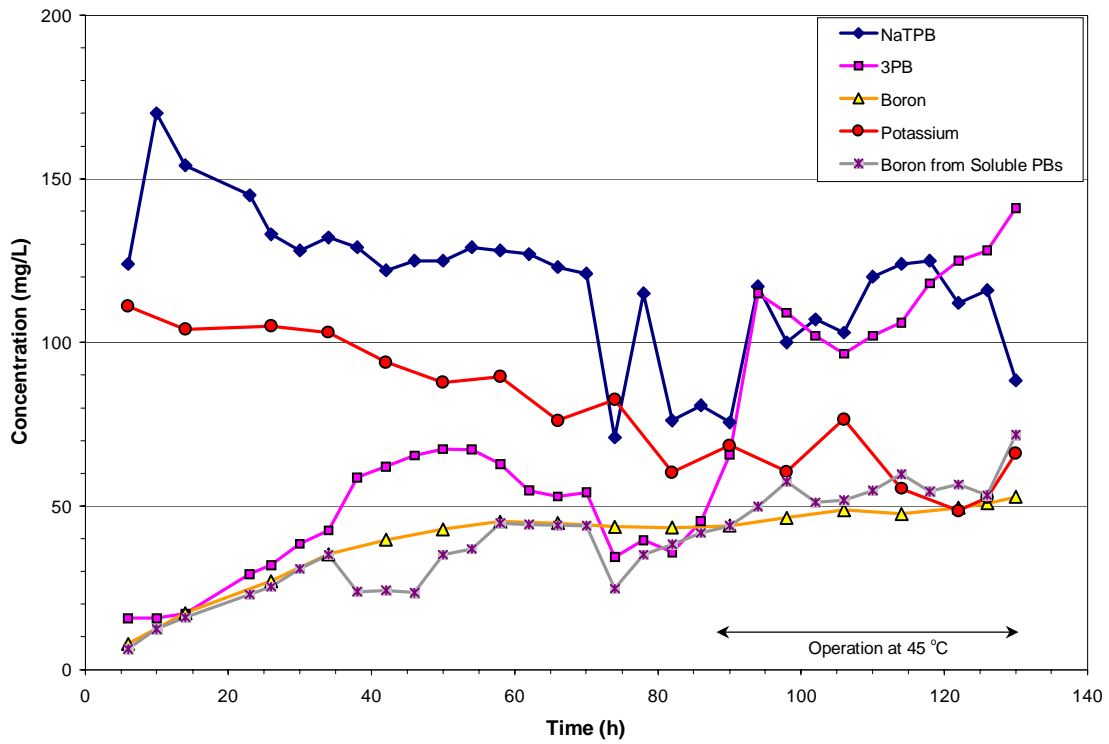


Fig. 22. Concentrate filtrate boron and potassium analysis results for Test 4.

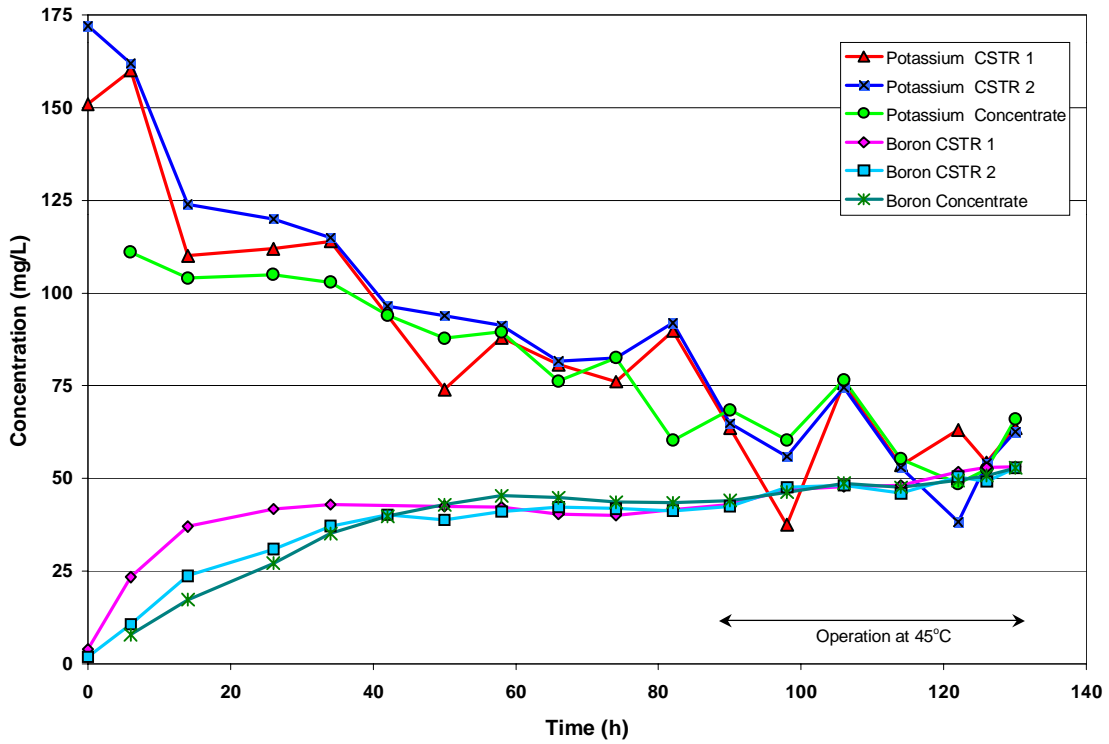


Fig. 23. Boron and potassium analysis results for Test 4.

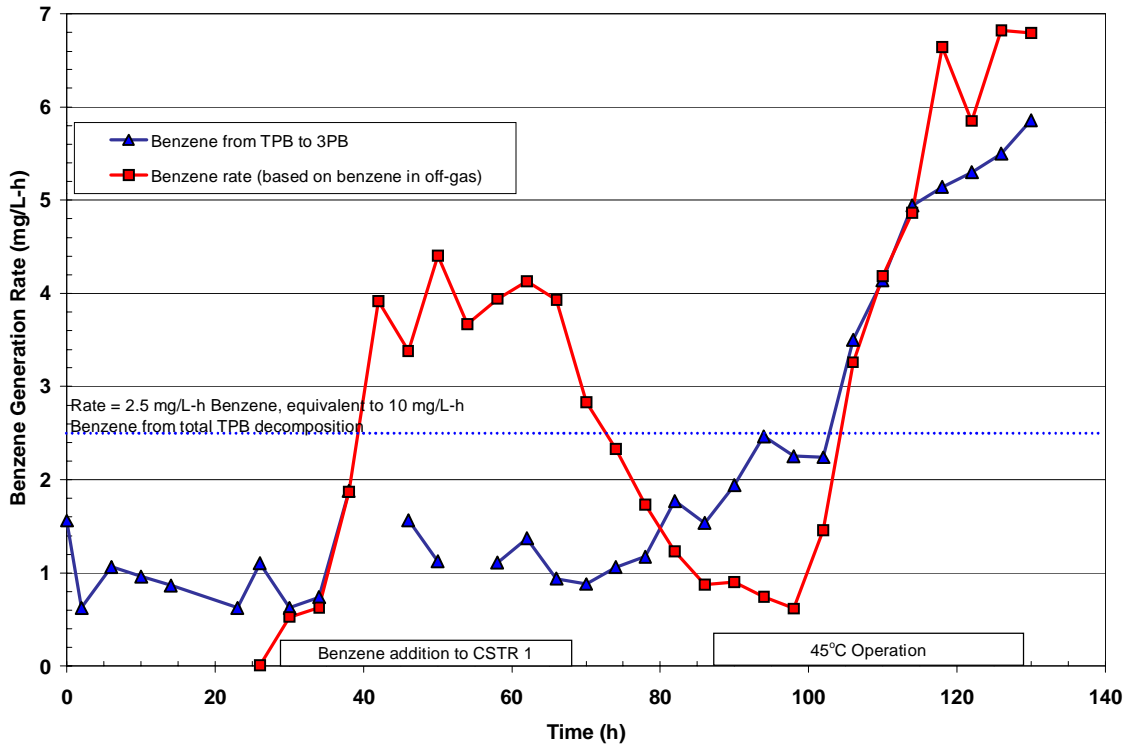


Fig. 24. Test 4 CSTR 1 benzene results.

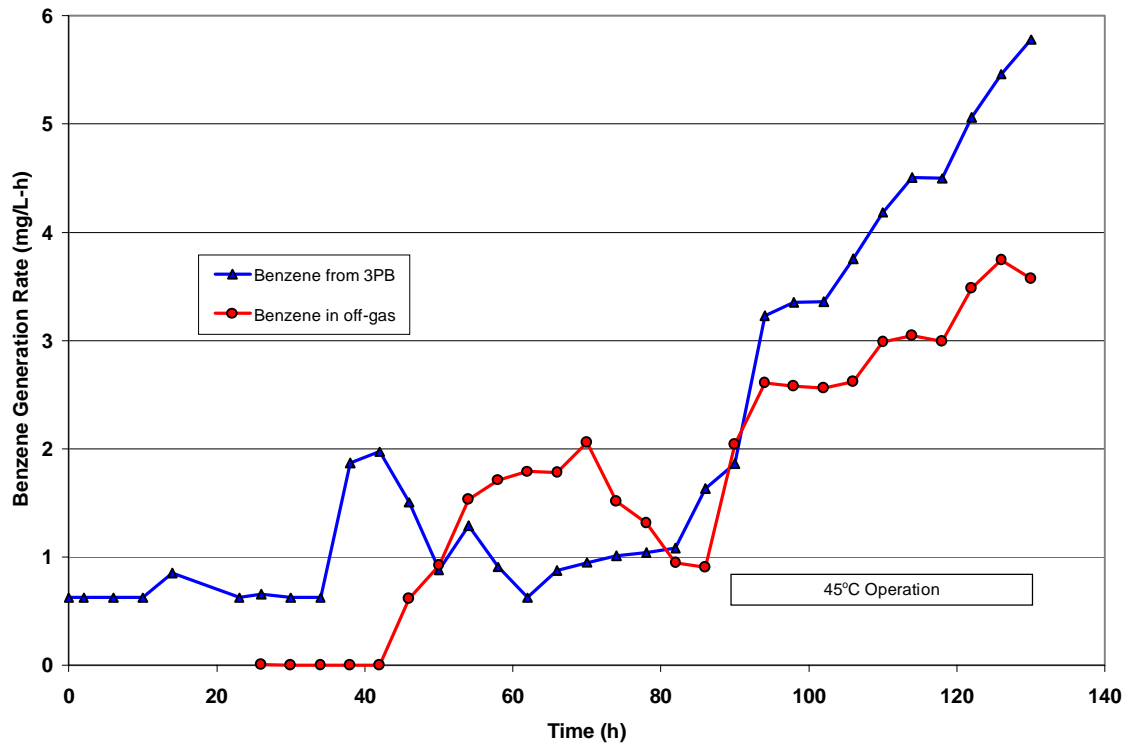


Fig. 25. Test 4 CSTR 2 benzene results.

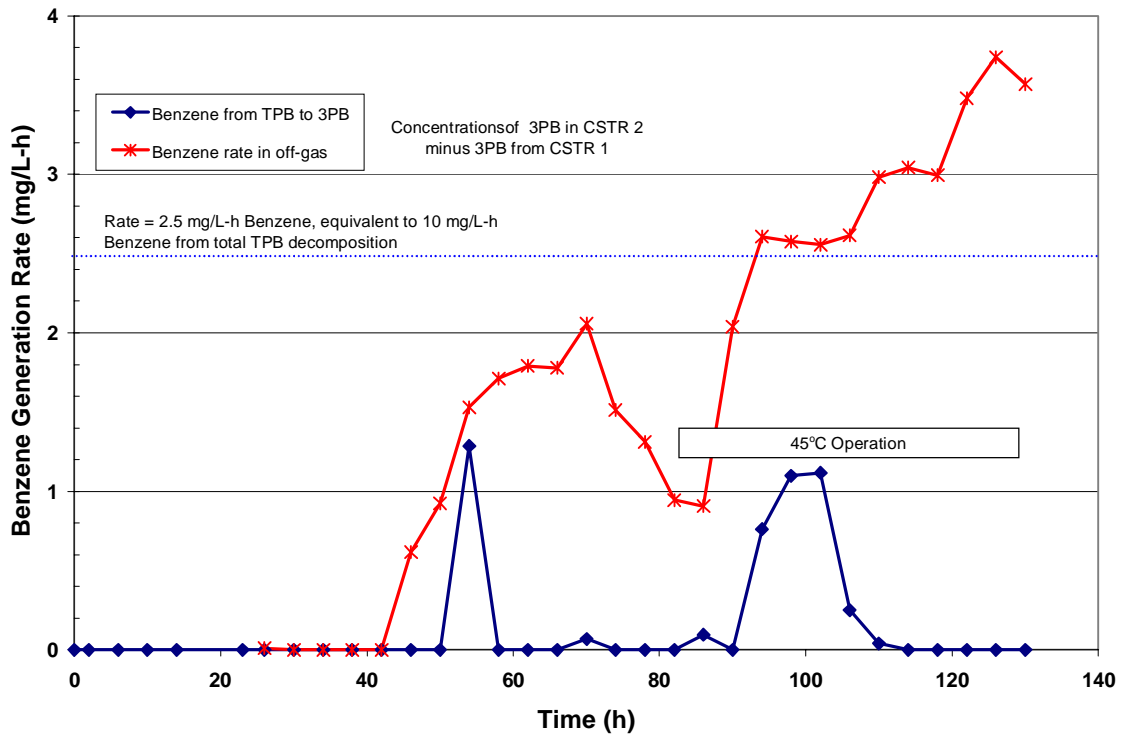


Fig. 26. CSTR 2 HPLC results for Test 4, subtracting CSTR 1 values from those in CSTR 2.

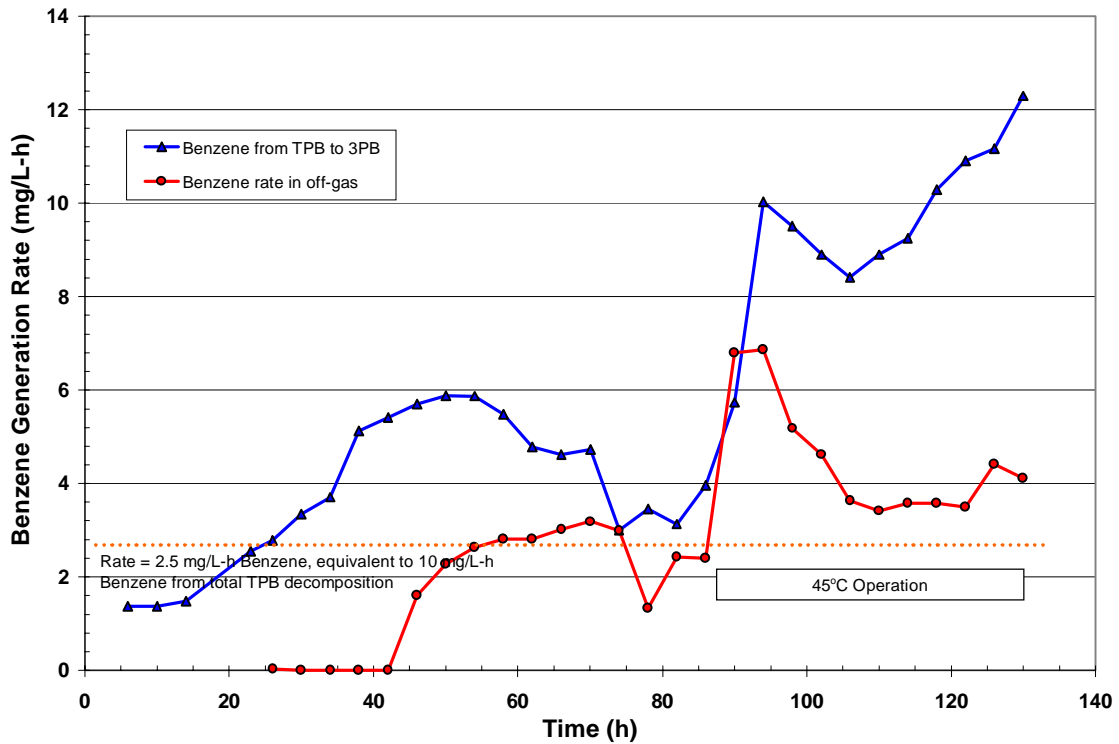
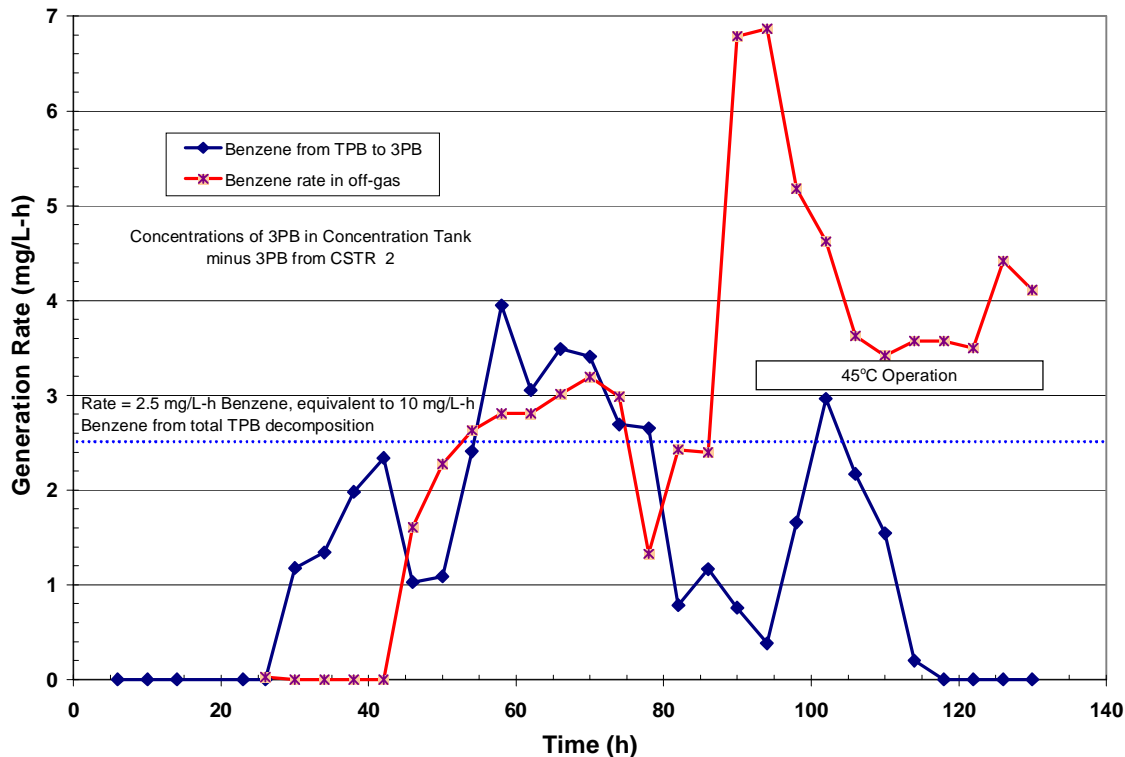


Fig. 27. Test 4 Slurry Concentrating Tank benzene results.



**Fig. 28. Concentrate filtrate benzene results for Test 4, subtracting CSTR 2 values from those in the filtrate.**

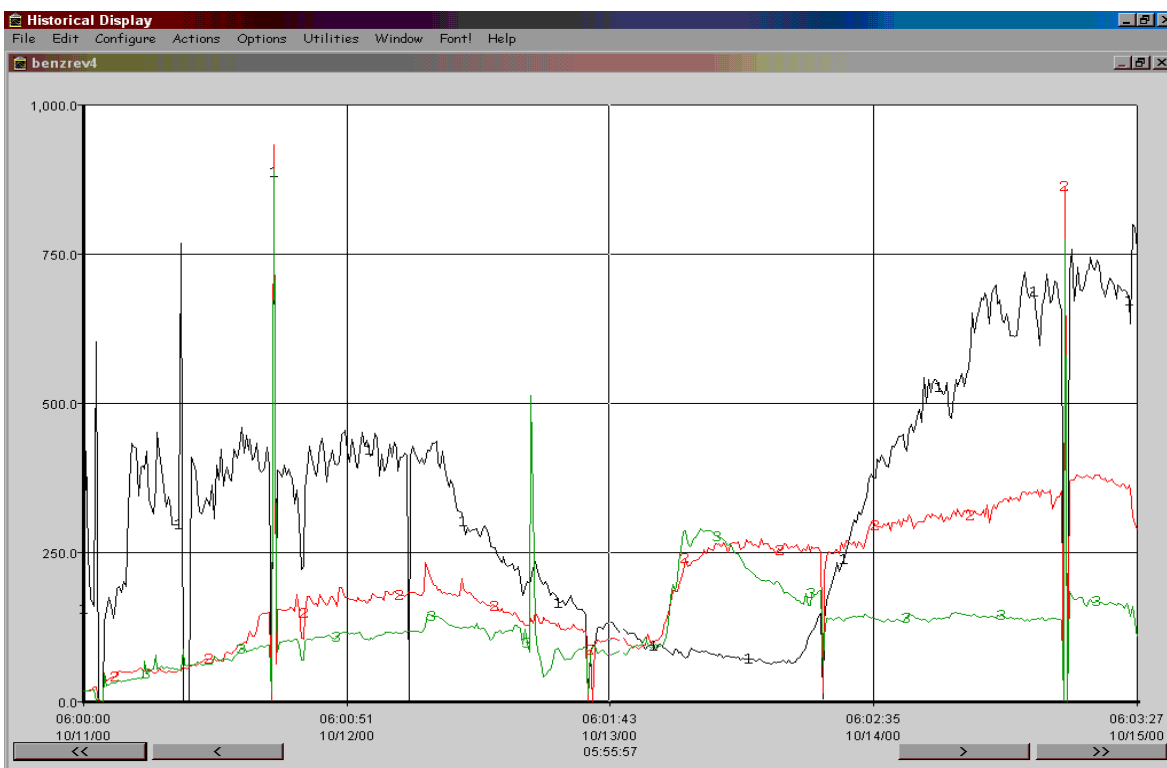
analysis of 3PB, the reaction of TPB with 3PB, producing 1 benzene molecule. As shown, there was a small generation rate (based on the production of 3PB from TPB), even before benzene was added to the tank. During the last 30 h of operation at 45°C, the rates based on reaction of TPB with 3PB and those based on the benzene in the off-gas closely agreed. Since the conditions that impact the solubility of benzene in the process fluid are not known, it is impossible to predict the expected concentration of benzene in the off-gas from the vessel at any particular time. One would anticipate, however, the presence of a measurable quantity of benzene in the off-gas and an overall increasing trend in the benzene content of the off-gas while the rate of TPB decomposition is increasing.

For CSTR 2, the benzene generation rates (Fig. 25) that are based on 3PB are similar to those obtained for CSTR 1. Benzene generation based on the 1PB concentration is not meaningful, because 1PB was artificially added as a separate feed. The calculated benzene generation rates based on the 3PB concentrations were significant during the first few hours, although no significant benzene was measured in the vessel off-gas until after about 50 h. The benzene rates based on the off-gas concentration after 50 h of operation compare closely with those calculated from the concentrations of 3PB in the filtrate samples. If it is assumed that the concentration of 3PB in CSTR 1 overflowing into CSTR 2 does not change, then that concentration can be subtracted from the CSTR 2 concentration to calculate the actual generation rate of the decomposition products in CSTR 2. The actual rate in CSTR 2 was significantly less than that calculated from the concentration of 3PB present. This is shown in Fig. 26, which indicates that the 3PB concentration was zero during most of the test. No 3PB was added to CSTR 1 via any of the supply feeds, and concentrations of 3PB at the start of the test were below the detection limit.

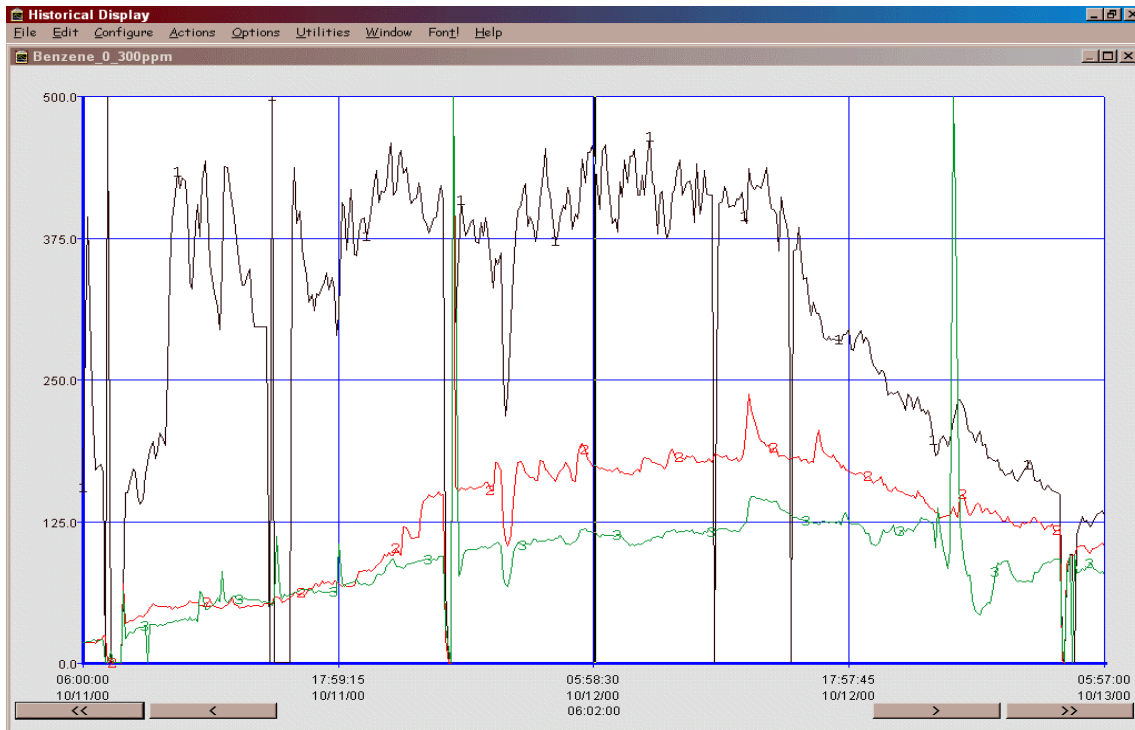
In the Slurry Concentrating Tank, the rates based on 3PB were similar to, but higher than, those in CSTR 2 (Fig. 27). The rates are calculated based on the smaller volume—and therefore faster throughput—in the Slurry Concentrating Tank versus the CSTRs. The rates based on 3PB present in the filtrate were significant within the first 24 h, although benzene production based on off-gas analysis was

not seen until after about 50 h. The benzene rates based on the off-gas concentration compare closely with those calculated from the concentrations of 3PB in the samples during the first batch of slurry at 25°C. If it is assumed that the concentrations of 3PB that enter the Slurry Concentrating Tank from CSTR 2 are subtracted from filtrate concentrations, then the rates are determined from the differences. When this is done, the benzene production rate in the Slurry Concentrating Tank is greatly reduced because some of the resultant concentrations are zero. Figure 28 shows that the production of benzene from 3PB in the Slurry Concentrating Tank is nearly zero during most of the test, indicating that most of the TPB decomposition to 3PB is occurring in CSTR 1.

The test results of the GASMET FTIR analyzers are shown in Figs. 29–31. Figure 29 shows an overall view of the 130 h of operation. The benzene concentration in the CSTR 1 off-gas declined after the benzene in the feed was stopped at 13:45 on 10/12/00 (or the 68th hour of the test). At 10:30 on 10/13/00 (the 88th hour), the temperature of the slurries reached 45°C. The increase in benzene generation is clearly visible for the remainder of the test. Figure 30 is an expanded view of the first part of the test at 25°C, and Fig. 31 shows the test while the slurries are being heated to and maintained at 45°C. The spikes represent times where the FTIRs were “zeroed” and calibrated. During the high-temperature portion of the test, the benzene concentration in the CSTR 1 off-gas was ~300 ppm higher than during the 25°C operation when benzene was being added to the CSTR with the TPB feed. Benzene off-gas concentrations for CSTR 2 were also higher during the high-temperature portion. The concentration of benzene in the Slurry Concentrating Tank off-gas rose slowly during the 25°C period. The concentration then remained fairly constant during high-temperature operation—after a peak when the slurry was transferred and the Slurry Concentrating Tank was refilled, an interval during which the slurry was not circulated through the filter system. These data clearly show that the TPB decomposition reaction was more active when the temperature was elevated.

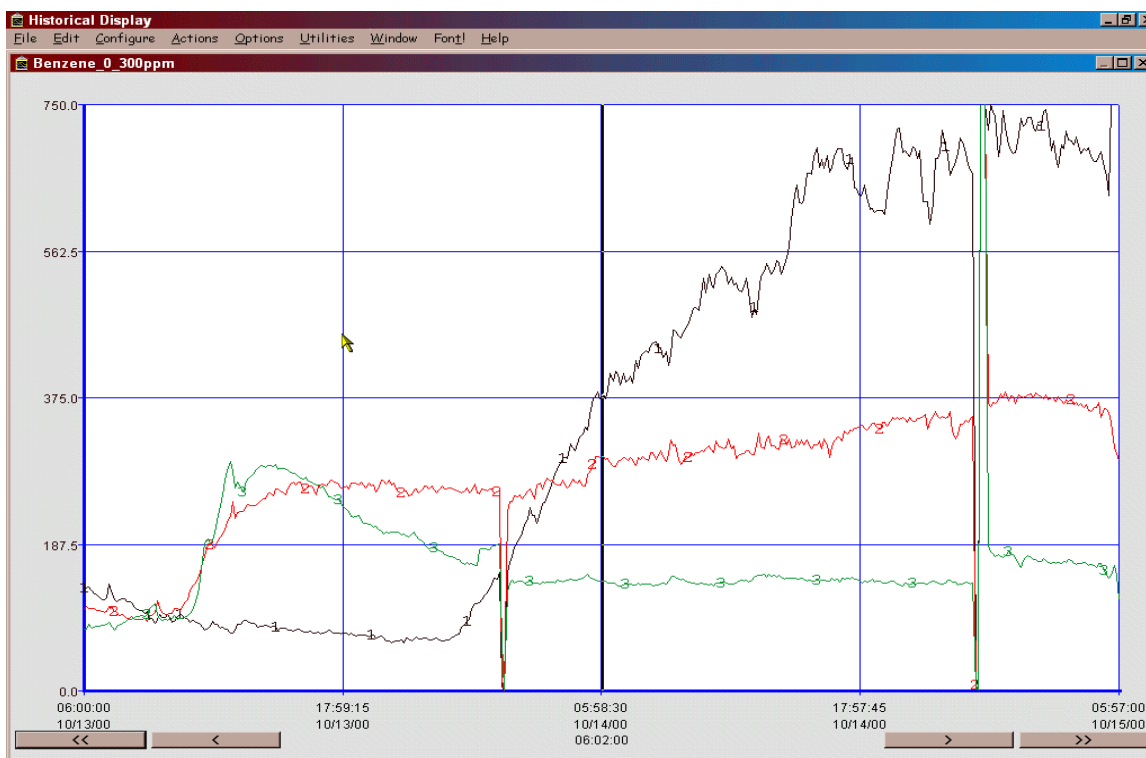


**Fig. 29. Test 4 benzene analyzer results.**  
(Four-day overview. Black—CSTR 1, Red—CSTR 2, Green—Slurry Concentrating Tank)



**Fig. 30. Test 4 benzene analyzer results.**

(Response during benzene addition. Brown—CSTR 1, Red—CSTR 2, Green—Slurry Concentrating Tank)



**Fig. 31. Test 4 benzene analyzer results.**

(Response during 45°C operation. Brown—CSTR 1, Red—CSTR 2, Green—Slurry Concentrating Tank)



#### 4.2.5 Material Balances

The feed supply streams to the CSTR system are shown in Figs. 32 and 33. Figure 32 shows the average flow rates for the several feeds. The flow averages were calculated only during the times when the feed systems were operational. Therefore, for the 5-h period that the system was down (hours 18–23), a horizontal line is shown on all of the feed inputs. The horizontal line readings for filtrate rate at 78 h (Fig. 33) resulted from the transfer of slurry to the Slurry Washing Tank and the time required to refill the Slurry Concentrating Tank. The wide variability in the filtrate flow rate results from the fluctuations that occur when the cell door is opened to remove samples or change filtrate collection tanks. The change in pressure inside the cell changed the level probe readings in the Slurry Concentrating Tank (a dp level probe) and caused the level controller to adjust the filtrate flow rate to compensate for the changed level reading. The flow interruption in the catalyst/MST graph at about 98 h resulted from the plugging of the catalyst/MST feed line. To continue flow, the line was connected to the NaTPB line.

Additional information on the feed to the system is shown in Fig. 33. It shows the cumulative totals of each feed stream during the course of the test. The salt feed plot shows an increase in feed rate during the final 12 h of the test, even though the feed pump speed had not been changed. This indicated that additional feed was being lost due to a leak in the feed tank. The filtrate curve shows the zero-flow point when the slurry was transferred to the Slurry Washing Tank and the Slurry Concentrating Tank was refilled at about 78–80 h of operation. All of the curves show that the feed stopped at hours 18–23. The loss of catalyst/MST feed during hours 98–101 shows up as the horizontal red series of points. Also, the feed required to fill the concentrate tank at the start of the test, about 5 kg, and the analytical samples of slurry from each CSTR and filtrate from the concentrate tank (collected every 4 h, at about 60 mL per sample) total about 7.3 kg. Adding them gives about 30.3 kg, which is in good agreement with what was fed to the system and what was removed.

Further information was obtained on the material balance for the catalyst fed to CSTR 1 during the test. Since two separate feed batches of catalyst/MST were prepared and used, separate material balances can be determined for the two parts of the test in which they were used. The liquid and solids in each bottle were separated by filtration, and samples of the filtrates were then analyzed for soluble mercury. The entire quantity of solids for each batch was then dissolved and analyzed for Hg, Pd, and Ti. Based on the analyses, the slurry feed rate was 2.547 g/min for batch 1 and 2.858 g/min for batch 2. The desired rate was 2.14 g/min, so the overall delivery rates were 119 and 134% of the targeted levels for batches 1 and 2, respectively.

Because of the uncertainty concerning the actual amounts of MST, Hg, and Pd delivered to the CSTR system, careful analyses of the remaining slurry solutions in both feed bottles were made. The first feed bottle, which was used for 56.7 h during the first part of the test, contained 0.775 g Pd, 6.10 g Hg, and 30.7 g MST (air-dried weight). By subtracting this amount from the initial quantity of catalyst in the bottle, it was found that 0.703 g Pd, 7.75 g Hg, and 64.1 g MST (air-dried weight) had been pumped to CSTR 1. On average, the delivery rates to CSTR 1 were 0.207 mg Pd/min, 2.28 mg Hg/min, and 18.8 mg MST/min. Comparing these rates with the pretest goals of 0.238 mg Pd/min, 2.24 mg Hg/min, and 15.28 mg MST/min, it was determined that about 87% of the amount of palladium needed for the 3X concentration in the slurry was delivered during this time period. MST feed bottle 2 was used during the last half of the test. A similar analysis found that the delivery rates to CSTR 1 were 0.149 mg Pd/min, 2.35 mg Hg/min, and 18.4 mg MST/min. For this period of the test, only 63% of the amount of palladium needed for the 3X concentration in the slurry was delivered. Although the palladium concentration in the slurry was less than the 3X concentration, it certainly was enough to provide TPB decomposition, especially at 45°C.

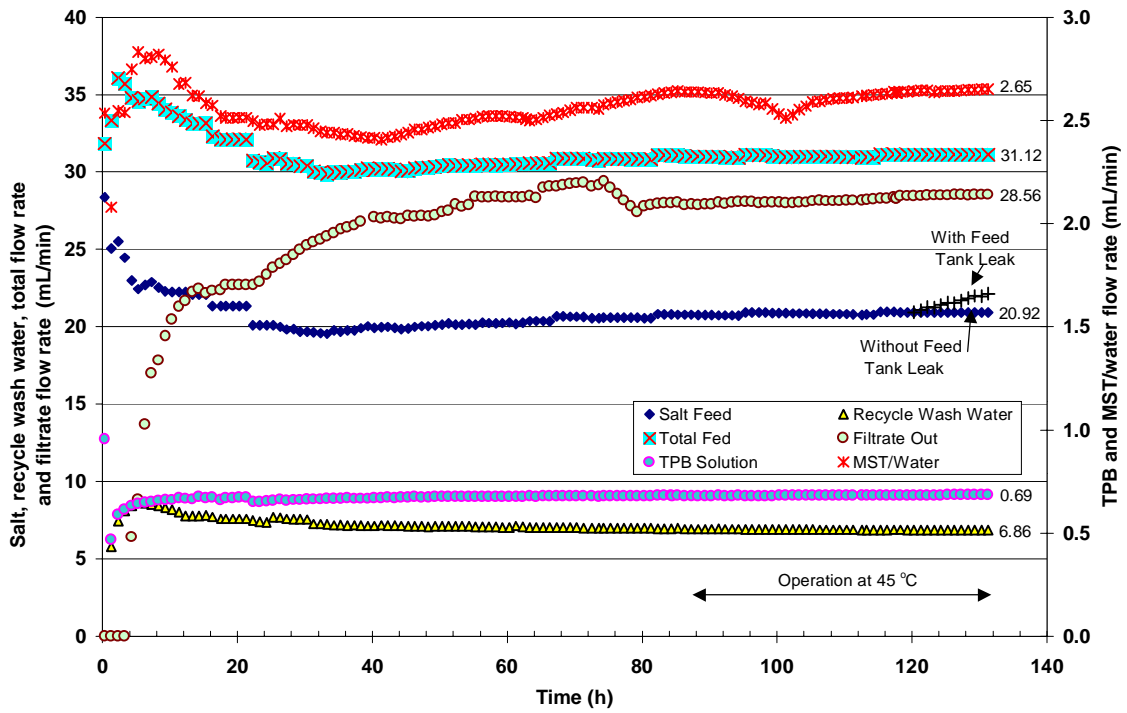


Fig. 32. Test 4 feed materials flow balance showing overall test average flow rate.

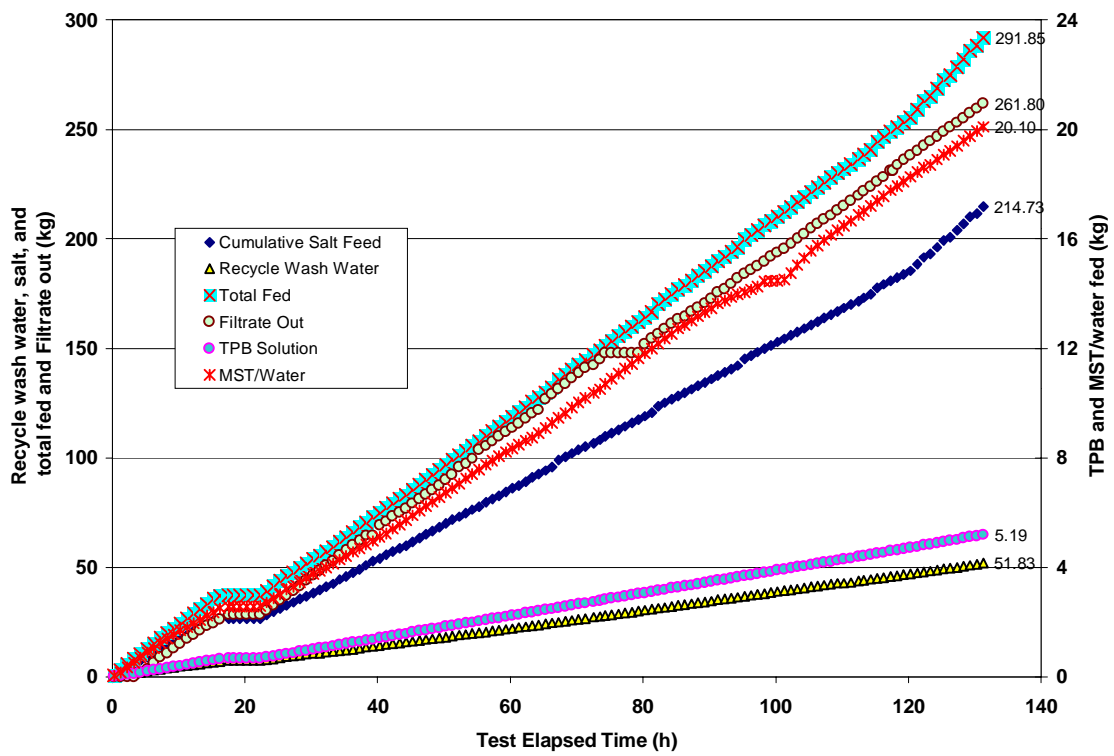


Fig. 33. Test 4 cumulative masses fed to CSTR 1.

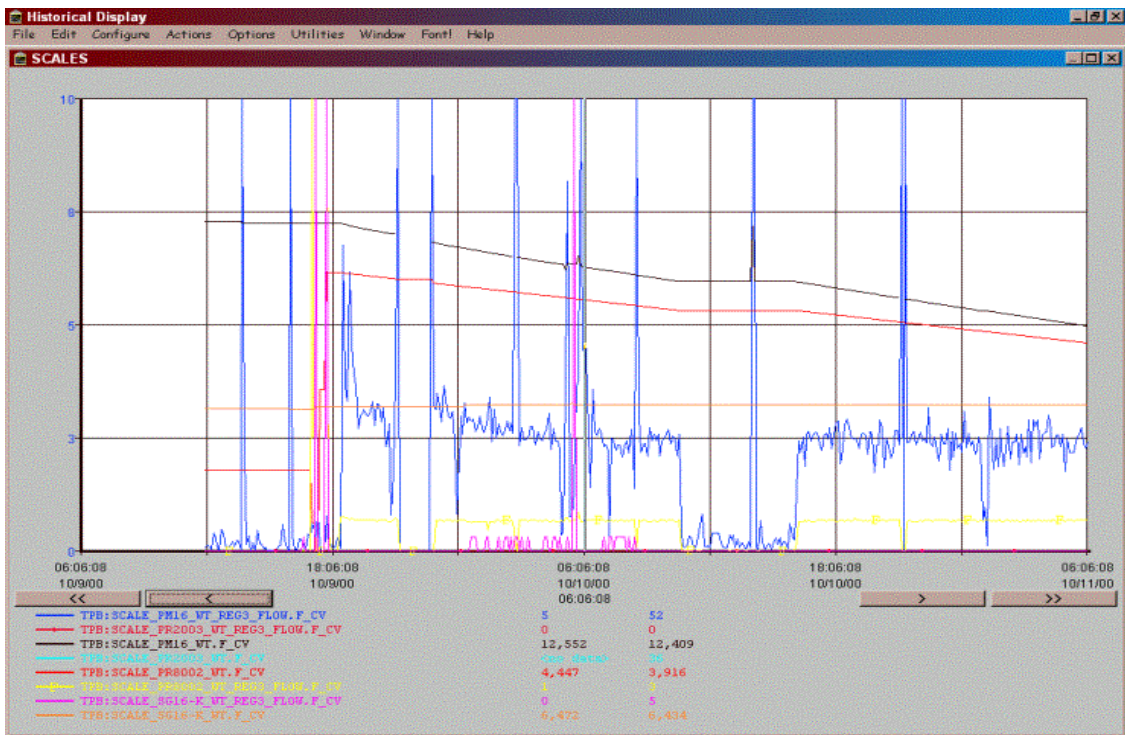
The results of the analyses, seen in Table 5, provide additional information for determining the feed rate of individual components of the catalyst/MST mixture. The mercury was fed at about 127% of the rate that was intended during both parts of the experiment. The amount of MST solids fed was about 124% of the planned amount for the first part and about 116% for the second part of the test. More palladium was fed during the first part due to the better suspension of particulates achieved in the slurry feed tank. During the second part, as noted above, the slurry tended to settle very quickly; made a large blob of sticky solids on the carboy bottom, which was very difficult to suspend; and tended to plug the feed line. It finally formed an unmovable plug at the check valve. The flow was stopped for several (4–5) h while the tubing was checked and the flow rerouted through the TPB line. The material balance results are shown in Table 6.

**Table 5. MST makeup for Test 4**

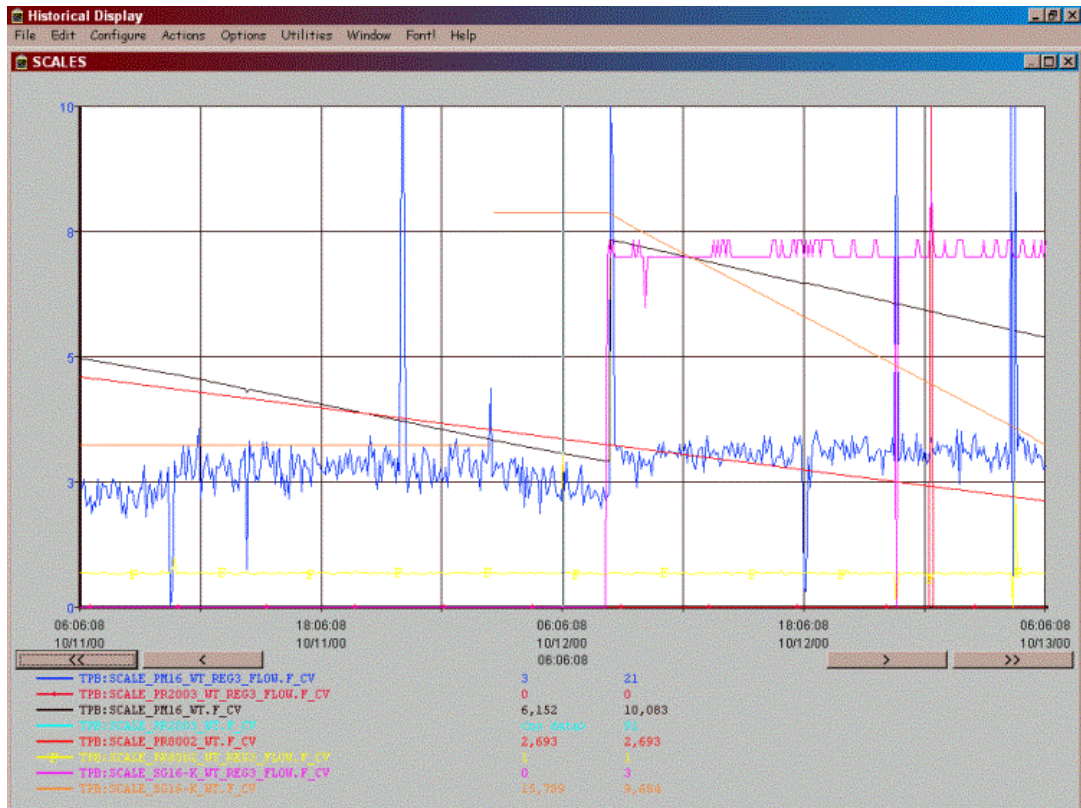
	Mass added to bottle (g)	Mass added (mg/g MST slurry)	Mass delivered (g)	Mass delivered (mg/g MST slurry)
<b>MST Bottle #1</b>				
Wet MST	220.23			
Air-dried MST	94.8	7.140	64.1	7.4
Hg(NO <sub>3</sub> ) <sub>2</sub> · H <sub>2</sub> O	27.61			
Elemental mercury	13.89	1.046	7.75	0.895
Palladium on alumina	29.56			
Elemental palladium	1.478	0.1113	0.703	0.0812
Water	13,000			
Total slurry #1	13,277.4		8,656	
<b>MST Bottle #2</b>				
Wet MST	270.8			
Air-dried MST	116.6	7.135	72.7	6.43
Hg(NO <sub>3</sub> ) <sub>2</sub> · H <sub>2</sub> O	34.33			
Elemental mercury	17.28	1.057	9.29	0.822
Palladium on alumina	36.33			
Elemental palladium	1.817	0.1112	0.5915	0.0523
Water	16,000			
Total slurry #2	16,341.7		11,305.1	

### 4.3 SUBSYSTEMS OPERATION

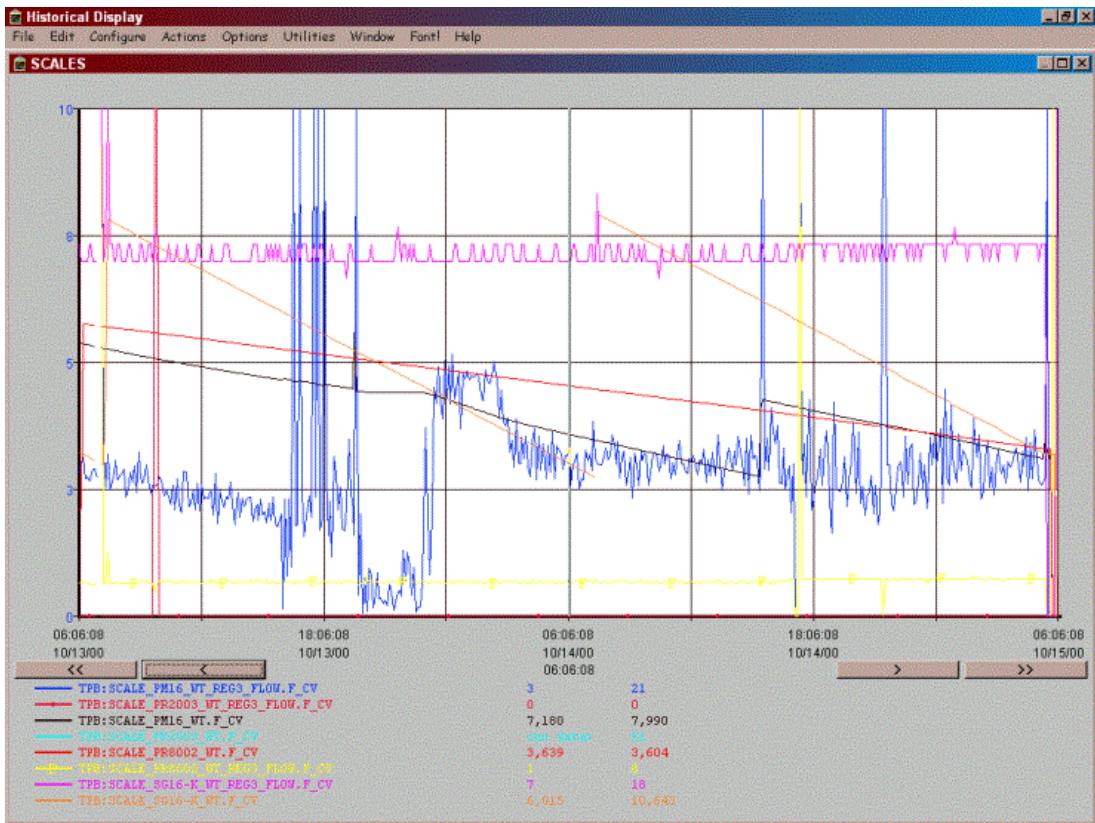
Figures 32 and 33 indicate that problems occurred with the catalyst/MST feed, especially during the latter part of the test. The plots in Figs. 34–37 from the computer data logging routine show more detail for the catalyst/MST stream and other feeds and indicate the times that the catalyst/MST flow stopped (shown by a slow decrease or sudden drop to zero flow). When flow declined below an established set point, the computer system sent an alarm that allowed the operators to begin to fix the problem immediately. Figure 36 shows the attempts to increase the slowing flow and restart the flow after it had stopped. Once flow had been restarted by feeding along with the TPB, the feed rate was set higher for several hours to compensate for the amount of catalyst/MST that should have been fed up to that time. Once the flow was restarted, no more stoppages occurred during the remainder of the test. No problems were noted with any of the other feed streams during Test 4, either in maintaining flow rate or with feed stoppages. Figure 37 shows the operational period during the test when recycle wash water from the Test 3 slurry washing was used.



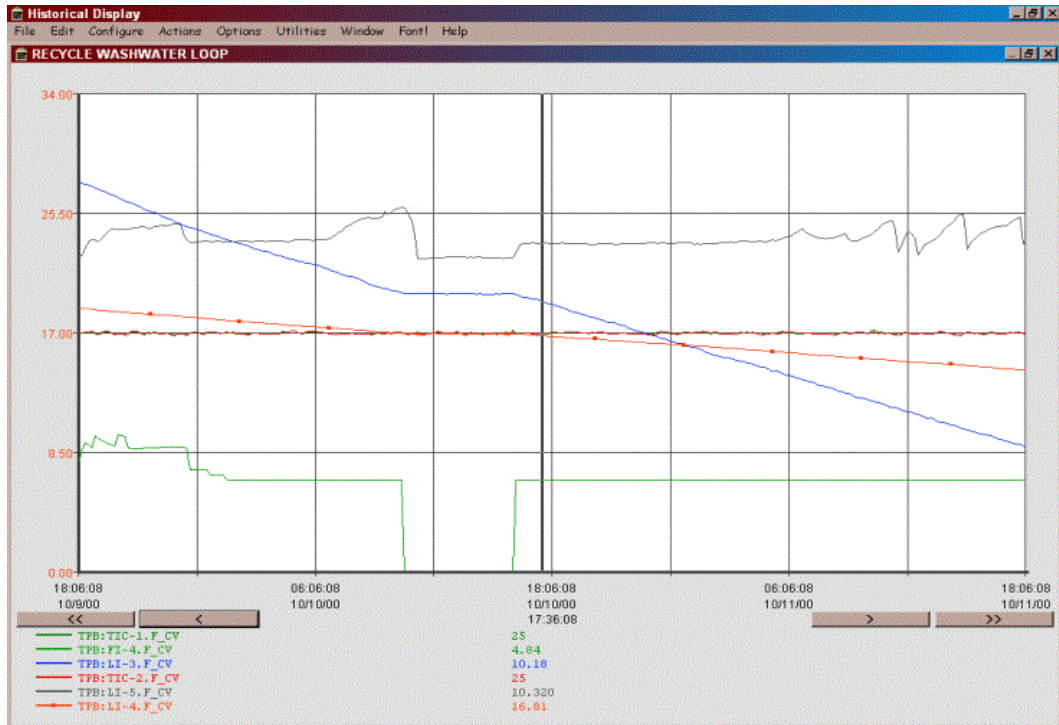
**Fig. 34. TPB and MST feed rates from their balances during 25°C operation.**  
 (Blue—MST/Water rate, Black—MST/Catalyst mass, Red—TPB mass, Yellow—TPB rate)



**Fig. 35. TPB and MST feed rates from their balances during heatup and 45°C operation.**  
 (Blue—MST/Water rate, Black—MST/Catalyst mass, Red—TPB mass, Yellow—TPB rate, Purple—Recycle wash-water flow, Orange—Recycle wash-water mass)



**Fig. 36. TPB and MST feed rates from their balances during 45°C operation.**  
 (Blue—MST/Water rate, Black—MST/Catalyst mass, Red—TPB mass, Yellow—TPB rate, Purple—Recycle wash-water flow, Orange—Recycle wash-water mass)



**Fig. 37. CSTR 1 operation during feeding from recycle washer tank.**  
 (Orange—Salt feed tank level, 0–34 in.; Black—CSTR 1 level, 0–15 in.; Blue—Recycle wash-water tank level, 0–18 in.)

**Table 6. MST slurry material balance for Test 4  
(Delivery to CSTR 1)**

	Actual/targeted (mg/L slurry)	Ratio (%)
MST Bottle #1 <sup>a</sup>		
Mercury	72.4/85	85
Palladium <sup>b</sup>	6.58/7.8	84
MST	597/490	122
MST Bottle #2 <sup>c</sup>		
Mercury	74.0/85	87
Palladium <sup>d</sup>	4.76/7.8	61.1
MST	578/490	118

<sup>a</sup>Actual slurry delivery rate—2.547 g/min; targeted rate—2.14 g/min; actual/target—119%.

<sup>b</sup>Targeted amount of palladium needed—3X; amount of palladium actually delivered—2.52X.

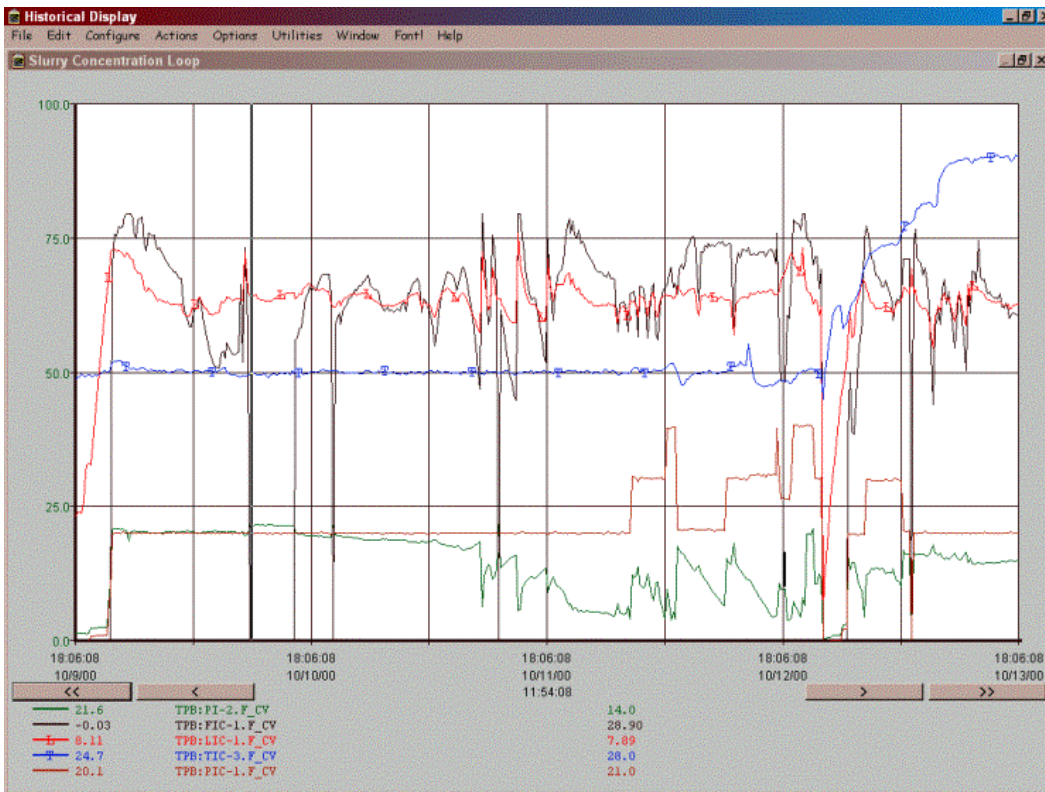
<sup>c</sup>Actual slurry delivery rate—2.858 g/min; targeted rate—2.14 g/min; actual/target—133.5%.

<sup>d</sup>Targeted amount of palladium needed—3X; amount of palladium actually delivered—1.83X.

Figure 37 also highlights the other hydraulic problem that occurred during the test when some form of solid material partially plugged the overflow outlet from CSTR 1 to CSTR 2 and caused the CSTR 1 level to rise. When this occurred the first time and flow could not be quickly reestablished, some of the CSTR 1 slurry was pumped directly to the concentrate tank using the peristaltic pump. This pump is normally used for waste removal purposes to process the slurry remaining in the CSTRs after a test is finished. This allowed time for the system feed pumps and the slurry concentration system to be shut down and a new valve to be installed on CSTR 1 in place of the plugged valve. As Fig. 37 shows, smooth flow was established for about 12 h and the valve then started to plug again. Tapping on the valve at this time and subsequent plugging times and tilting the valve in the direction of CSTR 2 seemed to allow the plug material to pass on through the valve. This procedure appeared to work for the remainder of the test, and the level in CSTR 1 remained constant.

Figures 38 and 39 show the operation of the slurry concentration system during the first and second slurry batch concentrations. Figure 38 shows 25°C operation and the backpulses that were required when the pressure drop across the filter became too great. The first backpulse occurred during the third day of processing when the backpressure reached 40 psi and the concentration tank level could no longer be maintained. Only three backpulses were required during the course of the testing. Figure 40 shows the performance response of the filtration system after the backpulse. The third backpulse was conducted during the temperature ramp up to 45°C, and no additional backpulses were required during operation at 45°C. The pronounced fluctuations in the flow rate and tank level throughout the test occurred when the cell doors were opened every 4 h to remove the slurry samples and filtrate sample from the cell for analysis. The change in ambient pressure in the cell was detected by the level sensor in the Slurry Concentrating Tank, causing an unneeded response. This occurrence did not significantly impact the overall operation of the filter.

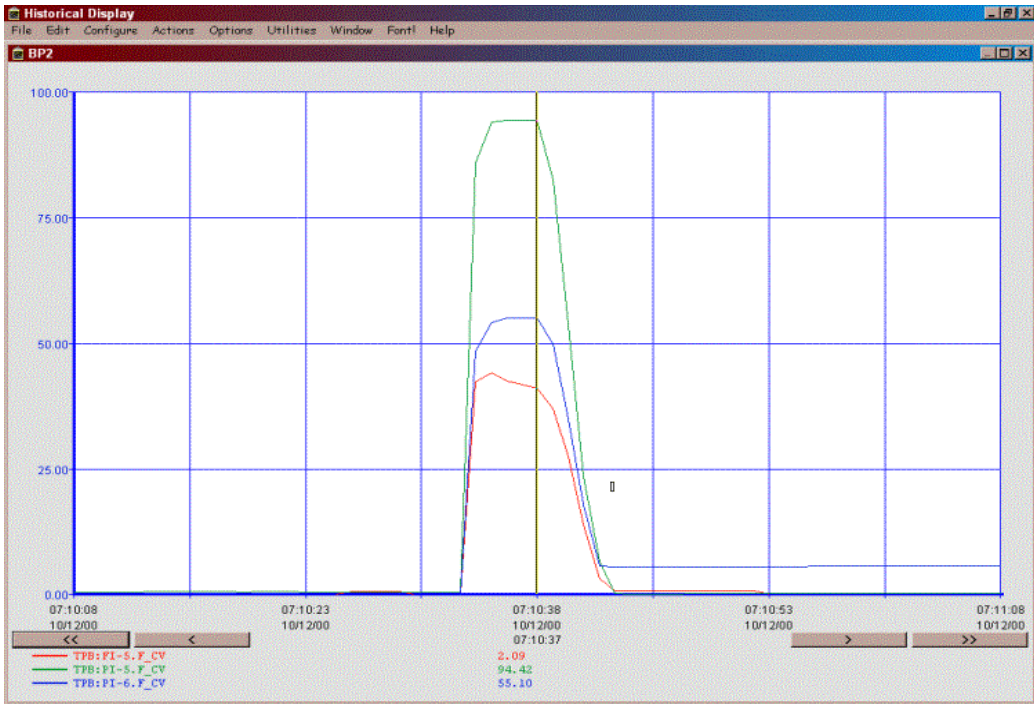
Figures 41 and 42 show the temperatures of the three process vessels during the 45°C operation. Also shown in the figures are the fluid temperatures in the temperature control circulators. The figures show the good temperature control achieved during this part of the test.



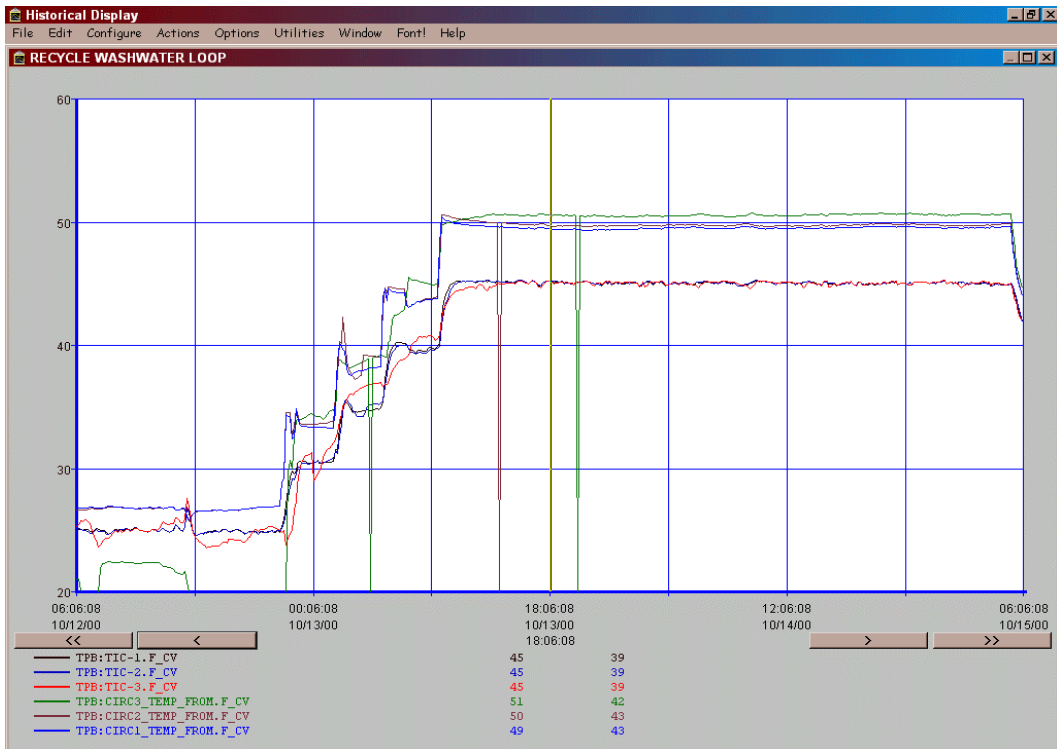
**Fig. 38. Slurry concentration loop operation during Batch 1 of Test 4.**  
 (Blue—Temperature, 0–50°C; Black—Filtrate flow, 0–50 mL/min; Red—Tank level, 0–11 in.;  
 Brown and green—pre- and post-filter pressure, 0–100 psi)



**Fig. 39. Slurry concentration loop operation during Batch 2 (45°C) of Test 4.**  
 (Blue—Temperature, 0–50°C; Brown and green—pre- and post-filter pressure, 0–100 psi;  
 Black—Filtrate flow, 0–50 mL/min; Red—Tank level, 0–11 in.)

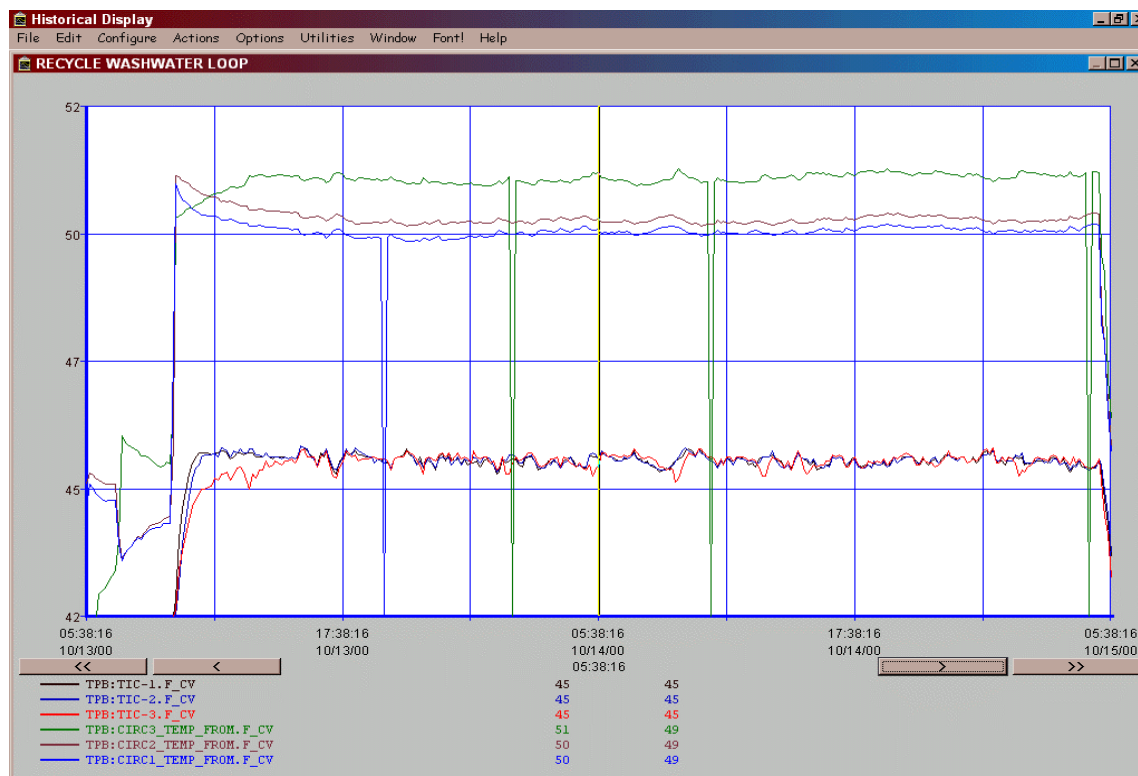


**Fig. 40. Slurry concentration loop filter backpulse during first batch concentration.**  
 (Green—Gas pressure, 0–100 psi; Blue—Loop pressure, 0–100 psi; Red—Liquid flow, 0–5 L/min)



**Fig. 41. CSTRs and Slurry Concentrating Tank temperatures during operation at 45°C.**  
 (Black—CSTR 1; Blue—CSTR 2; Red—Slurry Concentrating Tank)





**Fig. 42. Close-up view of CSTRs and Slurry Concentrating Tank temperatures during operation at 45°C. (Black—CSTR 1; Blue— CSTR 2; Red—Slurry Concentrating Tank)**

#### 4.4 SLURRY WASHING

The goal of washing the concentrated slurries is to remove, by dissolution in water, the excess TPB that was not precipitated with the potassium and cesium. The recovered TPB solution is then recycled to the CSTR 1 for enhanced TPB utilization. Both potassium tetraphenylborate (KTPB) and cesium tetraphenylborate (CsTPB) are insoluble under the dilute conditions; therefore, there is no threat of recycling additional dissolved cesium to the process. The extent of and desirability for recycling TPB will be determined by the end user of the process. Although Test 4 was not a closed-loop test, recycle wash water generated by washing the Test 3 concentrated slurry was used to start the test. A sufficient quantity of NaTPB was added to the wash water to provide the 60% excess TPB. In all previous tests in which the concentrated slurries were washed with water, the best effort to remove the TPB removed only about 30% of the needed amount. To provide the 60% excess, the TPB concentration in the wash water must be 0.033 mol/L.

The first batch of concentrated slurry in the Slurry Concentrating Tank (about 7 L at ~10 wt % solids) was transferred to the Slurry Wash Tank after 78 h of operation. About 1.5 L of concentrated slurry was left in the Slurry Concentrating Tank. The second batch was collected over the remaining 52 h of the test. At the end of the test, the slurries in both CSTRs were pumped from the bottom of the CSTRs to the Slurry Concentrating Tank and then concentrated. Because the recycle wash water tank was full after the process of washing the first batch of slurry, this concentrate batch was water washed in the slurry concentration filtration system and the wash water permeate from the filter was collected in carboys at the concentrator filtrate collection point. Both batches of slurry that were produced in Test 4 were washed with water in steps of 6 to 8 h for a total of about 24 h for use in Test 5. The washing rate was about 20 mL/min, and about 29 L of wash water was collected for each batch washed. Samples of wash-water

filtrates from the cross-flow filter were sent to RMAL for nitrite and TPB analyses. The results are given in Tables 7 and 8 and plotted in Fig. 43. The removal goal was to decrease the nitrite concentration in the wash-water filtrate samples to 0.01 M (440 mg/L). Tables 7 and 8 show that the nitrite concentrations in the last wash-water samples were 721 and 769 mg/L. The concentrations of nitrite and TPB in a sample of the combined wash water from the first batch were 5740 and 1120 mg/L, respectively. The TPB concentration (1120 mg/L) was only 15% of the targeted concentration of 7473 mg/L. These results were similar to those obtained for the Test 3 slurry wash. A sample of the composite wash from the second batch has not yet been taken for analysis.

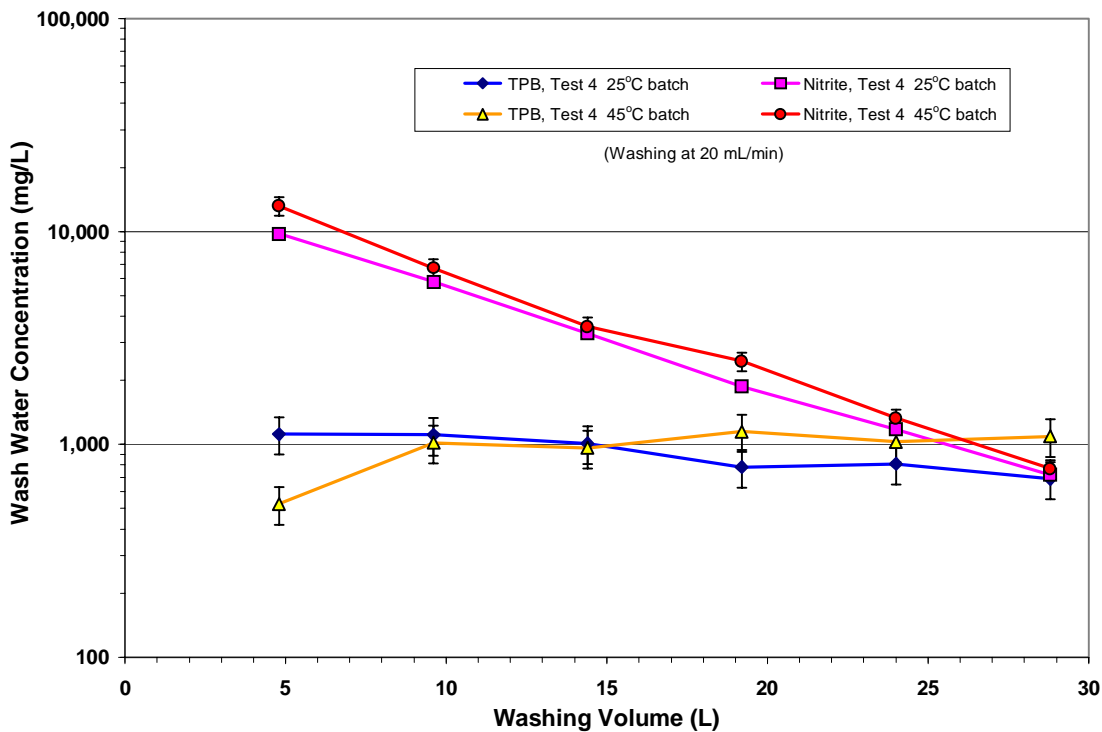
In this test and in previous tests, the expected quantity of TPB to be recovered from the wash water was not achieved. The solubilities of the TPB precipitates of Na, K, and Cs are the basis for use of the STTP.<sup>7,8</sup> The soluble-TPB data for the first batch of slurry concentrate that was washed decreased as the concentration of nitrite decreased. The sodium concentration in the wash filtrate should follow the same trend. For a nitrite concentration of 721 mg/L (0.0104 mol/L), the calculated Na<sup>+</sup> concentration would be 0.11 mol/L. If the precipitate had been pure NaTPB, it should have easily gone into solution at that Na<sup>+</sup> concentration. (The synthetic recycle wash water has a sodium concentration of about 1.2 M and a NaTPB concentration of 0.033 M, or about 11,300 ppm.) It appears that the solubility of NaTPB has been reduced due to some chemical condition or component of the concentrated slurry. In lab-scale tests at Savannah River, IITB52 proved to be the most effective antifoam/defoam agent; however, the test data indicated that TPB wash recovery was poor. For example, when no antifoam was present during washing, about 60% of the excess TPB was dissolved.<sup>9</sup> However, when IITB52 was used, only 12.7% of the excess TPB was recovered.

**Table 7. Results for Test 4 Batch 1 slurry washing at 20 mL/min**

Washing volume (L)	Washing time (h)	Wash tank filtrate sample ID	RMAL NO <sub>2</sub> <sup>-</sup> (mg/L)	RMAL NaTPB (mg/L)
4.8	4	T4W-1	9,740	1,120
9.6	8	T4W-2	5,810	1,110
14.4	12	T4W-3	3,330	1,010
19.2	16	T4W-4	1,870	784
24	20	T4W-5	1,180	810
28.8	24	T4W-6	721	689
Composite from recycle wash water tank				
29.500		36T3W-10	5740	1120

**Table 8. Results for Test 4 Batch 2 slurry washing at 20 mL/min**

Washing volume (L)	Washing time (h)	Conc. tank filtrate sample ID	RMAL NO <sub>2</sub> <sup>-</sup> (mg/L)	RMAL NaTPB (mg/L)
4.8	4	TW4H-1	13,200	526
9.6	8	TW4H-2	6,720	1,020
14.4	12	TW4H-3	3,580	965
19.2	16	TW4H-4	2,460	1,150
24	20	TW4H-5	1,330	1,030
28.8	24	TW4H-6	769	1,090



**Fig. 43. Nitrite and TPB concentrations during Test 4 slurry washing (25°C batch and 45°C batch).**

A mole accountability analysis of K, Cs, and TPB for each of the CSTR tests was performed to determine the available moles of TPB not associated with K and Cs in each batch of concentrated slurry that was washed (see Table 9). The analyses accounted for the moles of K, Cs, and TPB in the CSTRs and Slurry Concentrating Tank before each test and subsequently the moles of each constituent added during each test. It accounted for the moles of constituents that were transferred to the Slurry Washing Tank and the moles of each that were left in the concentrating system. In Test 1a, for example, all the slurry in the concentrating system and in both CSTRs was transferred to the Slurry Washing Tank and washed. In Test 1a, as well as Tests 2 and 3, the method of starting the test was to first fill the CSTRs with the 4.7 M Na<sup>+</sup> water diluted salt feed (an additional 3 L was added to the Slurry Concentrating Tank). Consequently, the initial batches of concentrated slurry had TPB/(potassium + cesium) mole ratios that were much lower than any subsequent batches prepared, as in Tests 2 and 4. Mole accounting for Test 1a found that the concentrated slurry contained 1.677 mol K<sup>+</sup>, 0.0124 mol Cs<sup>+</sup>, and 2.173 mol TPB. The available TPB (that was not precipitated by potassium and cesium) was 0.484 mol. The concentration of TPB in the slurry wash water was analytically determined by HPLC to be 0.0047 mol/L. The 12.3 L of wash water contained 0.058 mol TPB. Only 12% of the available TPB in the slurry concentrate was recovered. If all the available TPB had been recovered, its concentration in the wash water would have been 0.039 mol/L.

The slurries were left stirring in the CSTRs after Test 3 for use in the startup of Test 4. Only the concentrated slurry in the concentrating system was transferred to the Slurry Washing Tank. The K, Cs, and TPB analyses showed that 0.33 mol of NaTPB was available for dissolution and about 29% (0.095 mol) was recovered. At the start of Test 2, it was found that the recycle wash water from Test 1

did not have the needed TPB concentration to provide a 60% excess. To compensate, an extra TPB feed supply was added that was pumped into a tee connected to the MST/water supply line. The extra TPB was added at a rate of 0.40 mL/min and provided the 60% excess. The recycle wash water in the hot cell was added at 6.8 mL/min for the first 30 h of the test using about 12.4 L; its TPB concentration was 0.0047. For the next 26 h, simulated recycle wash water containing 0.033 M TPB was added. At 56 h testing, about 5 of the 7 L in the concentrating system was transferred to the Slurry Washing Tank for water washing. This batch contained about 0.86 mol of available TPB. The K<sup>+</sup> concentration in the salt feed was 0.0163 M rather than 0.0171 M because of the addition of the large volume of catalyst solution that was used in the test. The water-washing rate was 10 mL/min, and 18.6 L of recycle wash water was collected. The TPB concentration was only 0.0039 mol/L; only 9% (0.072 mol) was recovered. The mole accounting method estimates that 18, 19, and 20% were recovered in the next three batches. These results are twice as good as those for first batch—but still very low recovery.

Mole analysis of the washed batches in Test 4 found that TPB recovery was still inefficient, only 10 and 11% removal. The recoveries are similar to those for Test 1a and the first batch for Test 2. The mixing speed in Tests 1a, 2, and 4 was about 600 to 700 rpm. Antifoam B-52 was used in Test 4 but not in Tests 1a and 2. The best TPB recovery (29%) was obtained in Test 3 in which B-52 was used, but the mixing speed in the Slurry Washing Tank was 1200 to 1400 rpm (about twice as fast as in the other tests), which could have aided the improved recovery. Three of the batches of concentrated slurry in Test 2 removed twice as much of the available TPB as was removed in Test 1a, Test 4, and the first batch of Test 2. This seems to indicate that B-52 may have inhibited the recovery process. It is obvious that the recoveries in all these test were much less than desired, and additional research is need to resolve the problem.

**Table 9. NaTPB recovery in slurry washing process in CSTR tests**

Test	Wash cycle	Time (h)	Wash rate (mL/min)	Available TPB <sup>a</sup> (mol)	RWW <sup>b</sup> (L)	TPB recovered (mol)	TPB <sup>c</sup> (mol/L)	Percent recovered <sup>d</sup>
1a	1 <sup>e</sup>	16	10	0.48	12.3	0.058	0.0047	12
2	1	31	10	0.86	18.6	0.072	0.0039	9
2	2	28	15	0.63	25	0.115	0.0046	18
2	3	43	8	0.76	20.6	0.148	0.0072	19
2	4	44	5	0.69	13.2	0.138	0.0105	20
3	1 <sup>f</sup>	32	15	0.33	28.8	0.095	0.0033	29
4	1 <sup>g</sup>	30	20	1.09	35.4	0.117	0.0033	11
4	2 <sup>g</sup>	29	20	1.11	34.6	0.111	0.0032	10

<sup>a</sup>Moles of excess TPB as NaTPB in batch available to be dissolved by water washing.

<sup>b</sup>RWW = recycle wash water.

<sup>c</sup>Testing goal was to obtain enough TBP in recycle wash water to provide a 60% excess (~0.033 mol/L).

<sup>d</sup>(mol TPB recovered)/(mol available TPB in batch).

<sup>e</sup>Recycle wash water used in Test 2.

<sup>f</sup>Recycle wash water used in Test 4.

<sup>g</sup>Recycle wash water to be used in Test 5.

## 5. CONCLUSIONS

Test 4 was a 130.3-h test (actual operating time of 125.5 h) conducted to verify that acceptable STTP performance could be maintained while the TPB was actively decomposing. An improved catalyst mixture, which provided process concentrations of 7.8 mg/L Pd(0), 85 mg/L Hg(II), 500 mg/L phenylboronic acid (1PB), and 720 mg/L benzene, was added continuously to CSTR 1. The process temperature was maintained at 25°C for the first 76 h of operation while concentrating the slurry in the Slurry Concentrating Tank to about 10 wt %. Over the next 13 h, after the concentrated slurry had been transferred to the Slurry Washing Tank, the temperature of the slurries in the system vessels was increased in increments of 5°C to 45°C, where it was maintained for the remainder of the test. Decomposition of TPB was determined by HPLC analyses of the CSTRs and Slurry Concentrating Tank filtrates and by on-line FTIR measurement of benzene in the vapor space of each vessel. Foam suppressant IIT B52 was added to the system vessels to minimize foam formation.

The <sup>137</sup>Cs DFs for the filtrates from the Slurry Concentrating Tank were >56,000, with some values as high as 650,000. The decontamination chemistry exceeded the DF goal of 40,000 or 99.998% removal. Moreover, DFs for CSTR 1 and CSTR 2 filtrates were also very high, ranging from 21,000 to 140,000, with most being >30,000. The results indicated that the required cesium DF could be maintained despite the elevated temperature and high rate of TPB decomposition. The MST effectively removed the strontium, providing DFs of 100 to 200 for the duration of the test. The SRS WAC goal was 26, or 96% removal.

Efforts to remove the excess TPB by washing the produced slurries with water were not very successful. Only about 10–11% of the excess TPB that was calculated to be in the sodium form was removed. This is about the same percentage as was removed in Test 1a and the first batch of Test 2, in which the mixing speed in the Slurry Washing Tank was about 600 to 700 rpm. Based on these results, it does not appear that the use of IIT-B52 in Test 4 inhibited IIT-TPB recovery. Furthermore, in Test 3, B52 was used and three times more TPB was recovered than was recovered in Tests 1a, 2, and 4. However, a higher mixing speed (1200 to 1400 rpm) was used in Test 3, which may have enhanced the recovery. About 18 to 20% of the excess TPB was recovered in the last three of the four batches of concentrated slurry washed in Test 2. No B-52 was used in Test 2. The recovery data do not yet make it clear as to whether B-52 enhanced or slowed the NaTPB dissolution process during washing. It is obvious that the recoveries in all these test were much less than desirable and additional research is need to resolved the problem.

Problems with the operation of the equipment during the test were minimal, especially in the slurry concentration cross-flow filtration system, at both 25 and 45°C. Backpulsing the filter worked well when it was required. The benzene monitoring system also performed well, although minor program adjustments were required to remove some interference that was experienced during operation. The primary problems involved the feed systems, mainly the Catalyst/MST feed, which repeatedly plugged due to the Hg, MST, and Pd interactions and resulted in plugging of the feed line. This problem made it difficult to suspend the slurry mixture in the feed tank. The combined process fluids in CSTR 1 also presented plugging problems in the sample valve between CSTR 1 and CSTR 2, requiring the valve to be changed out during the first day of the test. However, these setbacks were resolved quickly and efficiently by the operating personnel without any significant impact on the overall process or the accuracy of the operating data.

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