A Characterization and Evaluation of Coal Liquefaction Process Streams

Quarterly Technical Progress Report July 1 through September 30,1996

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Section 1

EXECUTIVE SUMMARY

CHARACTERIZATION OF SAMPLES FROM HTI RUN ALC-1

CONSOL completed characterization of 64 samples from five run conditions of HTI Run ALC-1 (227-94), in which raw and cleaned (oil-agglomerated at low pH) Black Thunder Mine subbituminous coal was fed and processed using only dispersed catalysts in the liquefaction reactors.

- Extraction of THF-soluble resid from the pressure-filter cakes was more complete when agglomerates were fed, leaving only 5% or less THF solubles in the extracted cakes. When raw coal was fed, the extracted cakes contained **9-34%** THF solubles. HTI also observed improved filtration during the periods that agglomerates were fed. Improved operability, if verified by additional work, could be an economically significant benefit of coal cleaning by oil agglomeration at low pH. **An** apparently higher Mo addition rate (see fourth bullet) may have contributed to the benefits of using oil-agglomerated coal.
- Other stream sample characteristics changed when oil-agglomerated coal was fed in Conditions 2-4, relative to when uncleaned coal was fed in Conditions **1** and 5. The ash content of the 0-6 bottoms samples was lower when oil-agglomerates were fed. The THFsoluble **524 "C+** resid concentration in the feed **slurry** doubled when agglomerated coal was fed. Three factors may have influenced these characteristics. Higher coal conversion would have produced more resid. More efficient toluene-extraction of the filter cake would have recycled more resid. Removal of distillate as product to offset oil fed as part of the agglomerated coal would preferentially recycle the heaviest components.
- The analysis of process samples from Condition 5 indicates that the bench-unit dewaxing operations effectively removed paraffins from the vacuum distillate, and that the wax product is about 75% pure. The low concentration of wax in the feed minimized the impact of dewaxing on solvent quality improvement. The solvent hydrotreating operation effectively increased the hydroaromatic hydrogen content of the distillate and improved its solvent quality.

- CONSOL determined the concentrations of Mo and Fe in feed slurry and pressure filter cake ash samples from Run ALC-1. The Mo and Fe concentration results were combined with HTl material balance data to calculate Mo and Fe balances and determine the apparent addition rates of Mo and Fe catalysts. The resulting Mo and Fe material balance closures were between 80-120940. The results indicate that the Mo addition rate was inadvertently 10-169% higher than the target rate. The highest Mo addition rates were during agglomerated coal feed periods (13, 17, 20). The results show a low Fe addition rate in Period 13 (39% of the target rate), and **4%** to 36% higher than design rate in the other periods.
- \bullet Although the nominal fresh Fe and Mo catalyst addition rates decreased after Condition 2, there was an improvement in performance during Condition 3. The rate of Fe catalyst addition increased, based on the calculated Fe concentration from observed Fe concentrations in **two** streams. The observed performance improvement likely resulted from the increase in Fe catalyst concentration.
- Although HTI observed that the in-line hydrotreater (HTU) was less effective than expected in removing nitrogen and sulfur from the second-stage separator overheads (SOH), CONSOL data for all conditions show that this stream was effectively hydrogenated. It was hydrogenated to about the same degree **as** SOH produced in the coal-only operations during Run CMSL-9, and more hydrogenated than those produced in the coal-only operations during Run CMSL-11. The hydrotreated SOH oils were devoid of phenolic -OH, in spite of high phenolic -OH concentration in the unhydrotreated first-stage SOH oils, These observations suggest that the hydrogenation and phenol removal effectiveness **of** the catalyst was minimally affected by a distillation system upset early in Run ALC-1 that seriously decreased the N and S heteroatom removal ability of the HTU catalyst.
- Feed slurry samples were found to contain less than 5% 343 $^{\circ}$ C⁻ (650 $^{\circ}$ F⁻) material, contrary \bullet to concerns that large amounts of light oil were being recycled.
- About 1% of the feed coal carbon reports to the separator overhead water product. Error resulting from omitting this stream from elemental balance calculations should be small.

RECALIBRATION OF FTlR SPECTROSCOPIC METHOD FOR PHENOLIC -OH DETER-MINATION

The Fourier-transform infrared (FTIR) spectroscopic method used to determine the phenolic -OH concentration in liquefaction samples was recalibrated and implemented on a new FTlR instrument. **A** new calibration was obtained from spectra of standard phenol compounds. An error estimate for each determination was added to the analysis software. The method was validated by comparison of results from samples run on the old and new systems.

OIL ASSAYS OF HTI RUN PB-03 DISTILLATE PRODUCTS

CONSOL arranged to have crude oil assays conducted on the net products of HTI Run PB-03, at DOES request. Crude oil assays were conducted on net products obtained during periods of Run PB-03 for which the on-line hydrotreater was and was not used.

RESID REACTIVITY

- Resid reactivity tests were performed at CONSOL on the fifteen-resid sample set that the University of Delaware is using in their subcontract. The tests performed at CONSOL are described fully in the Results and Discussion section of this report.
- The University of Delaware made significant progress this quarter. Conversion values for all but two of the resids were obtained. The molecular structure model for coal-derived resids was assembled and is being optimized. A full description of the results is contained in the University of Delaware Quarterly Report appended to this report.

Section 2 INTRODUCTION

This is the Technical Progress Report for the ninth quarter of activities under DOE Contract No. DE-AC22-94PC93054. It covers the period July 1 through September 30, 1996.

CONTRACT OVERVIEW

The objectives of this project are to support the DOE direct coal liquefaction process development program and to improve the useful application of chemical analyses to direct coal liquefaction process development. This project builds on work performed during DOE Contract **No.** DE-AC22-89PC89883. Independent analyses by well-established methods are obtained for samples produced in direct coal liquefaction processes under evaluation by DOE. New analytical instruments and techniques to examine coal-derived samples are being evaluated. The data obtained from this study are used to guide process development and to develop an improved data base on coal and coal liquids properties. **A** sample bank, established and maintained for use in this project, is available for use by other researchers. The reactivity of the non-distillable resids toward hydrocracking at liquefaction conditions (i.e., resid reactivity) is being examined. From the literature and experimental data, a kinetic model of resid conversion is being constructed. Such a model will provide insights to improve process performance and the economics of direct coal liquefaction.

CONTRACT ACTIVITIES THIS PERIOD

- Characterization of samples from HTI Run ALC-1 was completed. Results are described in this report.
- The phenolic -OH measurement method was recalibrated for the new FTlR system (Appendix **1).**
- Sets of samples were requested from HTI Run PB-05 (Appendix 2).
- CONSOL arranged to have crude oil assays conducted on the net products of HTI \bullet Run PB-03, at DOE'S request. Crude oil assays were conducted on net products obtained

during periods of Run PB-03 in which the on-line hydrotreater was and was not in use. Results are provided in this report (Appendix 3).

- Resid reactivity tests were performed at CONSOL for a fifteen-resid sample set that the University of Delaware is investigating in their subcontract. The tests conducted at CONSOL are described in the Results and Discussion section of this report.
- The University of Delaware has made significant progress this quarter. Conversion values for all but two of the resids were obtained. The molecular structure model for coal-derived resids was assembled and is being optimized. A description of accomplishments is in the University of Delaware Quarterly Report appended to this report (Appendix **4).**

ACTIVITIES IN PROGRESS

- Characterization work was started on samples from HTI Run PB-04. In HTI Run PB-04, Black Thunder Mine coal, Hondo resid, auto shredder residue, and other plastics were used as feedstocks.
- Plans were made to begin a literature search and experimental work on the synthesis of ethyl phenyl ethers (for high-octane oxygenate extenders for transportation fuels) from coal liquefaction phenols. Reagents were ordered.
- The University of Delaware began a parametric study of resid conversion. Variables are time and temperature. The study will initially focus on the reactivity of two resids (Wilsonville Run 259, V131 B, and Wilsonville Run 260, V131 B).

Section 3

RESULTS AND DISCUSSION

CHARACTERIZATION OF SAMPLES FROM HTI RUN ALC-I

INTRODUCTION

CONSOL completed characterization of 64 samples from HTI Run ALC-1 (227-94), in which raw and cleaned Black Thunder Mine subbituminous coal was fed and processed using dispersed catalysts in the liquefaction reactors. The run consisted of 25 days of operation at five run conditions. Operating conditions and yields and performance data are provided in Table 1^{1-3} . The run background is described briefly below. Additional run information is presented elsewhere.¹⁻⁵

RUN OBJECTIVES AND HISTORY

There were four main objectives for Run ALC-1: 1) to provide a baseline operating period with Black Thunder Mine subbituminous coal, 2) to demonstrate liquefaction of low ash coal produced by low pH oil agglomeration of Black Thunder Mine coal, 3) to demonstrate liquefaction of Black Thunder Mine coal with dewaxing and hydrotreatment of distillate solvent, and **4)** to operate with extinction recycle of 343 *"C'* material, such that a hydrotreated light distillate would be the intended net product. An additional objective was to exploit the advantage of feeding low-ash coal agglomerates by decreasing the fresh catalyst make-up rate.

HTl's bench liquefaction Run ALC-1 consisted of 25 days of operation. Major accomplishments were:

- Oil agglomeration reduced the ash content *of* Black Thunder Mine coal by 40%, from 5.5% to 3.3% (MF, $SO₃$ -free ash basis).
- Excellent coal conversion (98%) was obtained with oil agglomerated coal (about 3% higher than the raw Black Thunder Mine coal), which increased the potential product yield by 2-3% on an MAF coal basis.
- Agglomerates were liquefied with no handling problems. Filtration performance was improved.

- Fresh catalyst (nominal) make-up rate was decreased by **30%,** with no apparent detriment to operations, when agglomerates were fed or when raw coal was fed (with solvent dewaxing and hydrotreating).
- Recycle solvent treatment by dewaxing and hydrotreating was demonstrated, but steady-state operation was not achieved.
- There was some success in achieving extinction recycle of the heaviest liquid products. This was evident in decreased resid yields and increased yields of light distillate during some run conditions.
- High yields of light distillates were obtained at several conditions during the run.

RUN PLAN AND OPERATING HISTORY FOR ALC-1

Run ALC-1 operating conditions, yields, and process performance results are shown in [Table 1.](#page-151-0) The fresh catalyst addition rates for Conditions 1 and 2 matched those used in Run CMSL-9, Condition 6. The fundamental recycle strategy planned for Run ALC-1 was to recycle to extinction all materials that boil above 343 "C, except for the rejected, washed, pressure-filter cake. In practice, equipment and operability constraints did not allow this. The intent was to adjust space velocity to extinct the yield of 343 *"c'* material. Pressure filtration was used to reject solids; the pressure-filter cake was washed with toluene to minimize rejection of solubles. The recycle sotvent (including buffers) to MF coal ratio was 1.60, and the recycle solids to MF coal ratio was 20%. Both of these are higher than typically used in recent runs at HTI. The run ALC-I conditions were designed to provide for high catalyst and unconverted coal recycling rates, which was intended to improve coal conversion, yet maintain operability with the high solids content in the recycle stream.

In Condition 1, the continuous atmospheric still bottoms (CASB) stream was filtered, and the recycle stream consisted of pressure-filter liquid (PFL) and a portion of the toluene-washed pressure-filter cake (PFC). The initial space velocity was 670 kg MF coal/h/m³. To shorten the solids recirculation time, it was decided near the end of Condition 1 to recycle unwashed PFC directly.

In Conditions 2, 3, and 4, oil agglomerated coal was used as the feedstock. The agglomerating oil used to prepare the ALC-1 feed was heavy recycle distillate (V1074) from the Wilsonville pilot plant Run 263 end-of-run inventory. For Conditions 2, 3, and **4,** it was necessary to reduce the quantity of distillate in the recycle by an amount equal to the agglomerating oil fed with the coal (as fed to liquefaction, the agglomerates contained about 23% oil). All of the PFL was sent to the vacuum still. **All** of the vacuum still bottoms (VSB) and some of the vacuum still overhead (VSOH) were recycled as necessary to meet total recycle requirements. A quantity of VSOH, approximately equal to the distillate agglomerating oil added with the coal, reported as product.

Early in Condition 2, the distillation conditions and flow scheme were changed to maintain operations with the constraints imposed by the run objectives. These changes resulted in a lighter recycle and heavier product than desired. The space velocity was decreased from 670 kg MF coal/h/m³ to 561 kg MF coal/h/m³ early in Condition 2. In Condition 3 (feeding agglomerates), the make-up catalyst rate was cut by 30%, relative to Conditions 1 and 2. Other distillation system changes were made to minimize recycle of light material. In Condition 4, the only operating condition changed was space velocity, which was decreased to 400 kg MF coal/h/m³ to reduce the yield of 343 *"C'* material.

The transition was made back to raw Black Thunder Mine coal in Condition 5. For this condition, which included dewaxing and hydrotreating, it was decided to keep the catalyst make-up rate at 70% of the Condition 1 value, and to set the space velocity at 481 kg MF coal/h/m³ reactor. The entire PFL stream was vacuum distilled at 524 "C to produce VSOH for dewaxing/hydrotreating. This material was ketone-dewaxed to remove paraffins, and the dewaxed oil then was hydrotreated to improve the donor solvent quality prior to recycle.

GENERAL SAMPLE CHARACTERISTICS

Sample characteristics are presented in Tables 2-6 and Figures 1-7. The aromatic hydrogen content of several whole liquid sample streams or filtered solids-containing whole sample streams is shown in Figure 4. Numerous changes made to distillation conditions and configurations during the run prevent direct comparison of vacuum distillate and pressure-filter liquid (PFL) streams. The streams shown in Figure 4 increase in aromaticity in the following order:

stage 2 separator overhead (SOH) oil < stage 1 SOH oil < feed slurry = interstage slurry < 0-6 bottoms = pressure-filter cake (PFC).

During the run, period-to-period variation in aromatic H content was generally small (Figure 4). However, for the feed slurry and interstage slurry streams, the Period 5-6 samples contained about **5%** to 8% more aromatic H than the Period 13-21 samples, which in turn contained about 5% more aromatic H than did the Period 25 samples. The aromatic H differences are probably associated with use of raw (Periods 1-6 and 21-25) vs. oil-agglomerated (Periods 7-20) coal, use of different space velocities in several run periods (space velocity was changed three times during the run), and use of recycle solvent dewaxing and hydrogenation in Periods 21-25. Analogous, but inverted, trends were observed for the corresponding paraffinic H content of these same samples (Figure **5).** The same pattern of high aromatic H content in the Period 5-6 samples, intermediate aromatic H content in the Period 13-20 samples, and lower aromatic H content in the Period 25 samples also holds for other sample types (Table 3), such as the IBP-454 °C (IBP-850 °F) distillate portion of the O-6 bottoms samples, and the 454-524 °C (650-975 **OF)** distillate portion of the feed slurry samples. The corresponding IBP-454 "C (IBP-850 °F) portion of the O-6 bottoms samples, and the 454-524 °C (650-975 °F) resid portion of the feed sluny samples show higher aromatic H in period 6, but nearly equivalent aromatic H for periods 13-25 (Table 4).

The phenolic -OH concentration of process streams (Table 5, Figure 7) decreased according to the trend:

first stage SOH > feed slurry = 0-6 bottoms = PFC > product SOH (second stage SOH)

No phenolic **-OH** was detected in the second-stage SOH samples. The phenolic -OH concentration in the feed slurry, 0-6 bottoms, and PFC streams decreased slightly as the run progressed. The phenolic -OH concentration in the first stage SOH increased after period 6, and then decreased for the remainder of the run. It is not clear why this particular trend is observed. The large increase after period 6 may be related to the increase in resid concentration in the feed sluny while agglomerates were fed (see later section). It **was** observed that the first stage SOH phenolic -OH concentration was inversely proportional to H_2 consumption on a MAF basis (% H_2)

consumption = -2.40*phenolic -OH concentration + 11.2, standard error of Y_{est} = 0.13, R^2 = 0.981 .

SOH OIL CHARACTERISTICS

As expected for normal hydrotreater operation, the second-stage SOH oils (the in-line hydrotreater product oils) are lower in aromatic H content and higher in paraffinic H content than the first-stage SOH oils. HTI indicated that the HTU was not operating at optimum performance during Run ALC-1, apparently because a distillation system upset early in the run caused some higher-boiling material to enter the HTU, decreasing the activity of the catalyst. Although HTI observed that the in-line hydrotreater was less effective than expected in removing nitrogen and sulfur from the second-stage SOH, CONSOL data (Table 3) show that the aromatic, hydroaromatic, and paraffinic hydrogen contents of these samples were generally similar to those produced in the coal-only operations during Run CMSL-9.6 The Run ALC-1 product (secondstage) SOHs also were less aromatic (more hydrogenated) than those produced in the coal-only operations during Run CMSL-11.⁷ These observations suggest that the hydrogenation effectiveness of the catalyst was minimally affected by the event early in Run ALC-1 that seriously decreased the N and S heteroatom removal ability of the HTU catalyst.

The hydrotreated product (second stage SOH) oils from HTI Run ALC-1 were devoid of phenolic -OH, in spite of high phenolic -OH concentration in the unhydrotreated first-stage SOH oils and of higher-than-normal concentrations of nitrogen and sulfur in the hydrotreated oil. The phenol removal activity, like the hydrogenation activity, evidently was not significantly impaired by the event that impaired the nitrogen and sulfur removal activity of the catalyst. The high phenolic -OH concentration of the unhydrotreated first stage SOH oil indicates that it contains a high concentration (perhaps 25%) of phenolic compounds, and is a potential source of phenolic chemicals.

ELEMENTAL ANALYSIS OF THE SOH WATER

Separator overhead (SOH) water samples from Stage 1 and Stage 2 were analyzed for pH and weight percent C, N, and **S** (Table 7) for HTl's elemental balances. Relative to the Stage 1 samples, the Stage 2 samples had higher pH and S and N concentrations, and lower C concentration. At the SOH water product flow rates in Run ALC-1 (about 14 kg/day), the carbon

content in the water is about 210 g/day (depending on the production split between the two stages). At the coal feed rate of Run ALC-1 (about 26 kglday), the coal carbon feed rate is about 16.6 kg/day. Based on this analysis, about 1 wt % of the coal carbon and somewhat greater proportions of the coal sulfur and nitrogen reported to the SOH water. It appears that any error resulting from omitting this stream from elemental balance calculations should be small.

The phenolic compounds in one SOH water sample (Period 138, first stage) were qualitatively characterized by GUMS analyses. The procedure used was acidification of the sample with HCI from a pH of about 9 to a pH of about 1 and extraction of the tar acids with methylene chloride. The appearance of the liquid changed from clear, bright yellow to colorless and slightly turbid upon acidification. Odors detected during acidification (in chronological order) included H_2S , SO_2 , and phenols. Only phenol and the cresols were identified by GC/MS analysis of the tar acids.

FEED SLURRY CHARACTERISTICS

There had been concern that significant amounts of light oils were recycled during Run ALC-1, due to inefficient distillation and other operating conditions. To address this concern, feed slurry samples from each condition of Run ALC-1 were distilled to 343 °C (650 °F) and 524 °C (975 **OF).** The resid was extracted with tetrahydrofuran, and ashed to determine other components as shown in [Table 8.](#page-35-0) An important result from this characterization is that only a small quantity of 343 °C⁻ (650 °F⁻) material was recycled in Run ALC-1 (\leq 5% of feed slurry). This presumes that no light material was lost before the samples were shipped to CONSOL.

CONSOL data (Figure 1, Tables **2** and 8) indicate that the feed slurry during Conditions **2-4** (Periods 13, 17 and 20) contained about twice as much THF-soluble 524° C⁺ (975 $^{\circ}$ F⁺) resid as the feed slurry from Conditions 1 and 5 (Periods 6 and **25).** The resid content of the Period 5 feed slurry was very high, for no apparent reason. Perhaps the higher concentration of higherboiling material in the feed slurry contributed to the lower distillate yields obtained when oilagglomerated coal was fed. Three factors may relate to the higher resid concentration in the feed slurry when oil-agglomerated coal was fed in Conditions 2-4. First, the higher coal conversion obtained when agglomerates were fed could result in a higher concentration of resid in the recycle streams. Second, removal of an amount of distillate roughly equivalent to that fed with the agglomerated coal may have left only the higher-boiling material for recycle. In fact, the recycle strategy for the entire run was to take the lowest-boiling material as product; perhaps this degree

of selectivity only became available with operating-condition changes that came into effect when oil-agglomerated coal was fed. The third possible factor is more efficient extraction of soluble resid from the PFC when oil agglomerates were fed (discussed below). This would increase the availability of soluble resid for recycle, although the quantitative effect on feed slurry composition was not considered. The recycle stream composition was controlled by the total recycle rate and solids recycle rate; soluble-resid recycle rate was not directly controlled by the plant operators.

CONCENTRATIONS OF Mo AND Fe IN FEED AND PRODUCT STREAMS

Mo and Fe concentrations in the second-stage pressure-filter cake (PFC) samples from Run **ALC-1** were determined. They are presented in [Table 9](#page-35-0) on three bases: (1) metal oxide as weight percent of the SO_3 -containing ash; (2) element as a weight percent of the SO_3 -free ash; and **(3)** element as a weight percent of the SO,-containing ash. The CONSOL Mo concentrations agree with those reported by HTI for the second-stage pressure-filter cake. However, many of CONSOL's Fe concentrations are 20% (or more) higher than those reported by HTI.

Mo and Fe concentrations were determined in the 524 $^{\circ}$ C⁺ (975 $^{\circ}$ F⁺) resids from all of the Run ALC-1 feed slurry samples (Table IO). The data provide a direct measurement of the total Mo and Fe concentrations in the reactor feed. The concentrations of the metals in the ash agree well with the pressure filter cake (PFC) results reported in [Table 9.](#page-35-0) Elemental Mo concentrations for the entire set (PFC and feed slurry) ranged from 0.19 to 0.47 wt % of the $SO₃$ -free ash, and approximately doubled when agglomerated coal was fed (partly in response to the reduced ash content of the agglomerated coal). Fe concentrations for the entire set ranged from 17.9 to 28.1 wt % of the SO,-free ash, and decreased from Condition 2 through the end *of* the run.

The Mo and Fe concentrations were combined with HTI material balance data to calculate Mo and Fe balances and to determine the apparent Mo and Fe catalyst addition rates. Intermediate and final calculated results are shown in Table 11. The following assumptions were used in the calculations: raw coal SO_3 -free ash content of 5.50% MF; agglomerated coal SO $_3$ free ash content of 3.30% MF; added $Fe₂O₃$ was 1.4% or 0.98% of the MF coal rate (for calculating total ash fed to system); Mo in the raw or agglomerated coal was ignored (the MF raw or agglomerated coal contains ca. 2 ppm Mo); the Fe₂O₃ content of the raw coal SO₃-free ash was 6.66%; and the Fe₂O₃ content of the agglomerated coal SO₃-free ash was 9.57%. The data on which the calculations are based were obtained from this report and other reports. $1,4-5$

The Mo and Fe balances were calculated for each test period shown as the ratio of total grams of Mo or Fe out to the total grams of Mo or Fe in. Total Mo or Fe in was obtained by multiplying the weight percent Mo or Fe in the feed slurry SO₃-free ash by the grams of ash fed in that period in the fresh coal and fresh iron catalyst, and in the pressure-filter cake (PFC) recycle. The amount of ash fed in the fresh coal was obtained from the amount of MF coal fed (HTI material balance data) and the $SO₃$ -free ash content of the raw or agglomerated coal. The ash fed in the added Fe catalyst was obtained as 1.4 wt % of the amount of MF coal fed (for the 10,000 mg/kg) Fe target addition rate), or as 0.98 wt % of the amount of MF coal fed (for the 7,000 mg/kg Fe target addition rate). The ash fed in the recycle PFC was obtained from HTl's recycle ratio of insoluble organic matter (IOM) plus $SO₃$ -free ash to MF coal, and HTI's measurements of IOM and SO,-free ash contents of the PFC. The total amount of Mo or Fe out was obtained by multiplying the wt % Mo or Fe in the PFC by the grams of total PFC produced for the period (HTI material balance data).

Apparent catalyst addition rates were obtained by subtracting the amount of Mo or Fe in the recycled PFC from the total amount of Mo or Fe fed in the feed slurry. The rates were calculated on an MF coal basis. In the Fe case, the coal Fe contribution was backed out, so that the results represent only added Fe. The ca. 2 mg/kg Mo in the raw coal was not backed out.

The resulting Mo balance closures were 80-1 19%, and the Fe balance closures were 98-120%. The balances were generally good; that is, there was not a consistent or gross imbalance between feed and product rates. These balances generally verify the validity and consistency of the analytical measurements. The calculated apparent Mo catalyst addition rates (MF coal basis) were 110-269 mg/kg Mo when the target Mo concentration was 100 mg/kg, and 94-161 mg/kg when the target rate was 70 mg/kg Mo. The calculated apparent Fe catalyst addition rates (MF coal basis) were 3,928-13,642 mg/kg Fe when the target Fe concentration was 10,000 mg/kg, and 7,309-8,571 mg/kg when the target rate was 7,000 mg/kg Mo.

These results indicate that the Mo addition rate was 10-169% in excess of the target rate. The excess amount seemed to be largest during periods when agglomerated coal was fed. The generally good Mo balance suggests that the Mo addition rate was inadvertently high, rather than being caused by introduction of Mo from a spurious source. Any spurious source of Mo would have to introduce a high Mo concentration in the feed slurry to be consistent with these results.

The results show that the Fe added in Period 13 (Condition **2)** was only 39% of the target rate, but in the other periods was 4-36% in excess of the target rate.

EFFECT ON PROCESS STREAM PROPERTIES OF FEEDING COAL CLEANED BY OIL AGGLOMERATION AT LOW pH

The ash content of the 0-6 bottoms stream was lower when agglomerates were fed in Conditions 2-4, Periods 13, 17, and 20 (9.3-9.8%, vs. 11.3-12.7% in Conditions 1 and 5). **A** similar but less pronounced pattern was observed in the feed slurry stream (6.9-7.8% vs. 8.2- 8.9% in Conditions **1** and 5). The lower ash concentrations in the samples seem to be related to the lower ash content of the agglomerated feed coal, but the relationship is not completely straightforward, especially for the 0-6 bottoms samples. Although the feed coal ash content was lower in the oil-agglomerates, the addition rate of Fe-based catalyst **was** changed, which contributes significantly to the ash content of the total feed. There is also a significant rate of ash recycle, which is larger than the feed rate of ash from the fresh coal. The ash recycle rate was held roughly constant, although changes that result from feeding agglomerates (such as an increase in coal conversion) would be expected to affect the recycle ash rate.

Extraction of THF-soluble 524°C⁺ (975 °F⁺) resid from the PFCs was more complete when oilagglomerated coal was fed in Conditions **2-4** (Periods 13, 17, and 20, Figure **3).** When raw coal was fed (Periods 6 and 25), 9-34% solubles remained in the toluene-extracted PFCs. When oilagglomerated coal was fed, only 5% or less THF-solubles remained in the toluene-extracted PFCs. HTl's data for hot-quinoline solubles confirm this effect.' Perhaps the improvement in extraction efficiency was related to the improved filtration observed when agglomerates were fed.⁸ If it can be demonstrated that oil agglomeration at low pH improves liquefaction plant operability, **low** pH agglomeration would have a significant economic impact. The apparently higher **Mo** addition rate during these periods may have contributed to the benefits of feeding oilagglomerated coal.

EFFECT OF THE DECREASE IN NOMINAL FRESH CATALYST ADDITION RATE

Decreasing the nominal fresh catalyst addition rate after Condition 2 had no detrimental effect on process sample characteristics. In fact, performance improved (as manifested by increased light distillate yield and decreased resid yield). **As** noted in a prior section, the concentrations of Mo and Fe in feed slurry and PFC samples suggest that the actual catalyst feed rates differed considerably from the nominal rates in some cases. Those results (Table 11) indicate that the Fe addition rate approximately doubled from Condition **2** to Condition 3, whereas the Mo addition rate decreased by 60% (but remained much higher than the target rate of 70 mg/kg). Perhaps the increased Fe rate contributed to improved performance during Condition 3. **No** operating condition obviously changed sufficiently to account for the improved performance, although changes in distillation configuration and distillation conditions could have some effect on performance that **is** difficult to account for.

EFFECT OF RECYCLE SOLVENT DEWAXING AND HYDROGENATION

Vacuum distillates from Run ALC-1 were dewaxed in the laboratory and the dewaxing feeds and fractions were analyzed. Microautoclave solvent quality tests and ' H-NMR spectroscopy were used) to determine the effectiveness of the Run ALC-1 dewaxing and hydrotreating operations. Either vacuum still overhead (VSOH) samples or the laboratory-generated 454 °C⁻ (850 °F⁻) distillate of 0-6 bottoms samples were tested.

[Table](#page-38-0) **12** shows that the 0-6 bottoms distillate contained about *56%* wax before dewaxing operations were started (Period 20). During dewaxing operations (Period **25),** the wax content of the second-stage heavy distillate (VSOH) product was reduced to 1 -2%, presumably mostly freshly produced wax. The heavy distillate (VSOH, dewaxed), also taken from period 25 but downstream of dewaxing, was virtually devoid of wax.

The 'H-NMR proton distributions and microautoclave solvent quality test results for heavy distillates relevant to the dewaxing operations are shown in Tables 3 and 6. The data are summarized in [Table 13;](#page-39-0) averages and standard deviations are shown where possible. The data in [Table 13](#page-39-0) show that the paraffinic nature of the heavy distillate was reduced by dewaxing and there was a minor improvement in solvent quality, as measured by **two** types of microautoclave liquefaction tests. Hydrotreating was effective in producing hydroaromatics and in improving solvent quality.

For reasons not clear, the VSOH and 0-6 bottoms distillates from periods 13, 17, and **20** were more phenolic than the corresponding untreated period 25 samples [\(Table 13,](#page-39-0) 0.98 meq/g vs. **0.54** meq/g). Dewaxing had no impact on phenolic **-OH** concentration, but hydrotreating virtually eliminated the phenols (0.07 meq/g found in hydrotreated samples).

In summary, the results indicate that the bench-unit dewaxing operations were effective. However, the low concentration of wax in the feed limited the impact of the dewaxing on improving solvent quality. The solvent hydrotreating operation increased the hydroaromatic hydrogen content of the distillate and improved its solvent quality.

WAX YIELDS AND CHARACTERISTICS

The wax produced during Run ALC-1 period 25 was characterized by elemental, 'H-NMR, and GC/MS analyses [\(Table 14\).](#page-39-0) The wax was black and had a strong odor. The elemental analyses indicate that it contained over 13% hydrogen; however, four replicate elemental analyses gave poor repeatability (about 5% relative). It appears that the reported carbon content (74 wt %) is incorrectly low, because the 'H-NMR and GC/MS analyses did not indicate major concentrations of any components except alkanes. The wax was de-oiled via CONSOL's standard dewaxing procedure at -5 *"C* in acetone to indicate wax punty. This yielded 75% de-oiled wax, but also 2% THF insolubles [\(Table 14\).](#page-39-0) The de-oiled wax has less than 1% aromatic protons [\(Table 14\).](#page-39-0)

RECALIBRATION OF FTIR SPECTROSCOPIC METHOD FOR PHENOLIC -OH DETER-MINATION

The Fourier-transform infrared spectroscopic (FTIR) method used to determine phenolic **-OH** concentration in liquefaction samples was recalibrated and implemented on a new FTlR instrument. A calibration was obtained from spectra of six standard phenol compounds. An error estimate for each determination was added to CONSOL's analysis procedure software. The method was validated **by** comparison of results from samples run on both the old and new systems. A detailed discussion is presented with the results in Appendix 1.

CRUDE OIL ASSAYS OF NET PRODUCTS OF HTI RUN PB-03

CONSOL arranged to have crude oil assays conducted on the net products of HTI Run PB-03, at DOE'S request. Crude oil assays were conducted on net products obtained during periods of Run PB-03 with and without use of the on-line hydrotreater.

HTI provided CONSOL with fifteen individual separator overhead (SOH) samples from HTI Run PB-03. The samples were contained in 1 qt (0.95 L) plastic bottles. CONSOL prepared and blended these materials to produce two large samples for crude oil assay testing. The blended samples represent periods with (Periods 6, 7, and 8) and without (Periods 9, 10, and 11) the online hydrotreater in operation. Most of the eight individual samples of hydrotreated oil contained a water layer and a sediment (or emulsion) at the oil/water interface. The blends consisted of only the sediment-free oil phase, which was obtained by decanting the easily separated portion of the oil phase, then passing the remainder of the decant through a filter to retain any water or sediment. The water, sediment, and a small remaining oil layer were discarded. Although there was no visual evidence *of* water or sediment in the seven individual unhydrotreated samples, their dark color made it difficult to ascertain this; therefore, the same procedure was used for blending them. The blended materials weighed 19.96 kg and 17.64 kg, respectively. Each of the *two* blended materials was placed in *two* steel 5 gal (3.79 L) cans and shipped to lnchcape Testing Services Caleb Brett Laboratory in Houston, TX. Caleb Brett's reports on the crude oil assays appear as Appendix 3.

Caleb Brett returned all remaining samples and fractions of the two crudes to CONSOL after completion *of* the crude oil assays. One of the crudes and several of the fractions could not be returned to CONSOL because they were consumed during the crude oil assays. CONSOL determined the proton distributions by 'H-NMR and the phenolic -OH concentrations by FTlR on the returned materials. Those results appear in Table 15.

RESID REACTIVITY STUDIES

Knowledge of the chemistry of resids and understanding of resid reactivity is key to improving direct liquefaction process design and process economics. One approach to acquiring this knowledge is to correlate the chemical characteristics of the resids with the process conditions under which the resids were produced. This was examined in CONSOL's Contract DE-AC22- 89PC89883.¹² Another approach is to correlate the relative reactivities of the resids with their chemical characteristics. In order to acquire reactivity data on coal liquefaction resids, a standard empirical test for resid reactivity was developed by the University of Delaware (UOD) under subcontract to CONSOL. Distillation resids chosen by CONSOL for their representation of many different process conditions are being tested (Appendix 4).

To further study the reactivity of resids and their correlation with chemical characteristics, two models are being developed by the University of Delaware: a structural representation model and a reactivity model. Both models can be used for predictive purposes. Analytical data acquired by CONSOL and UOD on the resids being examined by UOD will be incorporated into the structural model. Conversion data and reaction product characteristics will be used to train the reactivity model. However, the University of Delaware resid reactivity tests are being carried out in a short time batch reactor (STBR) with very rapid heating and cooling times. All products of reaction are consumed in the resid conversion determinations. Therefore, there are insufficient reaction products to obtain chemical characterization of them for development of the reactivity model. Additionally, the product gases from the STBR are not analyzed. In order too produce sufficient sample for analysis of products (including gases), a series of resid reactivity tests were made in 45 mL microautoclaves.

These resid reactivity tests were carried out at CONSOL on the set of fifteen resids that currently are under study at the University of Delaware. Temperature and residence time were the same as those used in the STBR: 420 *"C* for 30 min residence time. The reactor charge in all tests was 4.00 g resid, 8.00 g tetralin, and 10.3 MPa (1500 psi) (cold) H_2 . Each resid was reacted with and without molybdenum naphthenate catalyst precursor. In tests with molybdenum naphthenate, a charge of methyldisulfide (also known as dimethyldisulfide, DMDS) was added to sulfide the catalyst precursor. The molybdenum naphthenate charge, when used, was 2.00 g (3 wt % Mo on resid); each catalytic test used 0.5 g of DMDS. A full description of experimental methods is given in the Experimental section (Section 4).

Product analyses included gas chromatography of the collected gases, distillation of the condensed products to an atmospheric equivalent temperature of 454 °C (850 °F), and elemental analysis of the 454 °C⁺ (850 °F⁺) product fraction. Product distributions are presented in [Table 16.](#page-41-0) The table gives the feed, product distribution, and resid conversion on an SO₃-free ash-free basis. The analyses of the 454 °C⁺ (850 °F⁺) fraction of the products also are provided on that basis. Gas chromatography data were obtained on a volume percent basis. From the vent gas mass and the volume percent of each component, a gram quantity is determined; the gram quantities are provided in [Table 16.](#page-41-0) The nitrogen and oxygen concentrations are presumed to be due to air introduced during charging, sampling collection, and sample analysis; they are normalized out of the analyses. A blank space in [Table 16](#page-41-0) under Analysis of Product Gases indicates that GC analysis was not performed for that sample.

There was <<1 vol % methane and H₂S detected in the non-catalytic tests. In catalyzed tests, hydrogen reacts with the DMDS to produce methane and H_2S .

Many of the tests were repeated; these replicates are represented on [Table 16](#page-41-0) by a postscript letter (e.g., 22a means this was the second test at the same conditions of test 22, or 14c means that this is the fourth test at the same conditions as test 14). All analyses were sent to UOD for inclusion in the structural model.

A comparison to the resid conversion values obtained at Delaware [\(Table 17\)](#page-48-0) shows that the microreactor gave lower conversions than the STBR. The slow heat-up of the microreactor or other scaling factors may be responsible for the difference. Because of the differences in conversion values obtained with the *two* reaction systems, the microautoclave data were not included in UOD's modeling of resid conversion. The model could conceivably be reworked to account for different reactor systems and accept these data; however, there is no current plan to do this.

Section 4 EXPERIMENTAL

RESID REACTIVITY TESTS

The sequence for resid reactivity tests was: The total charge for all tests was 4.00 g resid, 8.00 g tetralin, and 10.3 MPa (1500 psi) (cold) H_2 . In tests with molybdenum naphthenate, a charge of methyldisulfide (also known as dimethyldisulfide, DMDS) was added to sulfide the catalyst precursor. The molybdenum naphthenate charge, when used, was 2.00 g (3 wt % Mo on resid); each catalytic test used 0.5 g of DMDS. After sealing, the vessel was connected to the shaker and shaking was started. While shaking, the microautoclave was lowered into a preheated sand bath. Reaction temperatures were achieved in two to four min. Reaction temperature for all tests was 420 °C. At the end of the 30 min residence time, the microautoclave was removed from the sand bath and immersed in ambient-temperature water while still shaking. The contents were quenched to ca. 100 *"C* in less than 90 sec. The 30 min reaction time includes the time from immersion of the microautoclave in the sand bath until immersion in the water bath.

A post-run weight was obtained to check for leaks. The microautoclave was attached to a pressure gauge at the fitting above the fill valve. It then was submersed in a dry ice/acetone bath until the internal thermocouple read -70 *"C.* The valve was opened and a cold temperature gas pressure reading was obtained. The microautoclave was removed from the dry ice/acetone bath and allowed to slowly warm to room temperature; pressure was continually monitored as the vessel warmed. When the internal thermocouple registered room temperature (20-22 "C), a final pressure reading was recorded. The gas was vented into a one-liter evacuated chamber; samples were taken from this chamber for gas chromatography analyses. Gas chromatography was performed using a four-column Carle Model Ill analytical gas chromatograph. A calibrated standard gas mixture [\(Table 18\)](#page-49-0) was used to calculate gas composition. The remaining gas was vented. The microautoclave and contents were reweighed. Gas make was determined by the difference in the weight before the tests and after venting.

Prior to distillation of the sample to 454 °C (850 °F), the microautoclave head was removed. A distillation head was attached. The microautoclave body was encased in a custom-designed heating mantle. The microautoclave was heated to a temperature of 265 °C (510 °F) at 5 torr (454 "C (850 **OF)** atmospheric equivalent temperature). The distillate was retained for further

analysis. The 454 °C⁺ (850 °F⁺) contents of the microautoclave were submitted for elemental analysis.

Resid conversion was determined on an SO_3 -free, ash-free basis as follows:

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resid conversion = 100 - \frac{filter\ cake - SO_3-free\ resid\ ash}{MAF\ resid} \times 100
$$

OTHER TESTS

The experimental procedures used to produce results presented in this report have been described previously.⁹⁻¹¹

Section 5

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HTI OPERATING CONDITIONS AND PROCESS PERFORMANCE RESULTS FROM RUN ALC-I (227-94)

Other Conditions:

1.60 Recycle to MF coal ratio (lowered slightly by wax yield in Condition 5).

0.20 Recycle solids to MF coal ratio.

15 MPa system pressure

Notes:

(a) Not including small quantities produced in DW-VSOH hydrotreating.

(b) Includes 4.79% wax product in 343-524 "C range.

(c) VSOH dewaxed and then hydrotreated (DWHT) prior to recycle.

COMPONENT DISTRIBUTION OF WHOLE SAMPLES HTI Run ALC-1

Note:

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[Feed slurry component concentrations are provided in Table 8](#page-12-0).

(a) Obtained by difference.

PROTON DISTRIBUTION OF WHOLE SOLIDS-FREE SAMPLES AND DISTILLATE SAMPLES HTI Run ALC-1

PROTON DISTRIBUTION OF FILTERED WHOLE SOLIDS-CONTAINING SAMPLES AND RESID SAMPLES HTI Run ALC-I

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PHENOLIC -OH CONCENTRATION IN SAMPLES FROM HTI RUN ALC-1

ND = Not Detected

MICROAUTOCLAVE CONVERSIONS HTI Run ALC-1

 (a)
 (b) Microautoclave coal conversion, wt % MAF, in modified equilibrium test, Old Ben Mine coal

(b) Pre-run make-up

(c) Distillate was 343 x 524 "C

[TABLE 7](#page-11-0)

COMPOSITION OF SOH WATER PRODUCTS FROM HTI RUN ALC-I

(a) Total S analysis by LECO SC-32 Sulfur Analyzer

(b) C and N analysis by LECO CHN-1000 Analyzer

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CONCENTRATION OF COMPONENTS OF THE FEED SLURRY SAMPLES FROM HTI RUN ALC-1

(a) Datelperiod in question

(b) May have contained an unknown amount of water

 (c) End point was ca. 529 °C (985 °F), instead of 524 °C (975 °F)

Actual distillation conditions were 5 torr/348 °F pot/378 °F head (176 °C pot/192 °C head) and 5 torr/608 °F pot/638 °F head (320 **"C** pot1337 **"C** head) to provide effective cut points of 343 **"C** (650 "F) and 524 "C (975 OF).

TABLE 9

ASH ELEMENTAL COMPOSITION **OF** PRESSURE-FILTER CAKES FROM HTI RUN ALC-1

Source: [Table](#page-41-0) **16** [\(page](#page-44-0) **40)** of the HTI draft Run ALC-1 report of June 1996 (early draft of Reference 1). (a)
ASH ELEMENTAL COMPOSITION OF FEED SLURRY 524 "C' RESIDS FROM HTI RUN ALC-1

(a) Single determination, others are averages of duplicate determinations.

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Mo AND Fe BALANCE AND APPARENT ADDITION RATES FOR HTI RUN ALC-1

ppm values are **on** a ppm **(mg/kg)** MF coal basis.

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WAX CONTENT IN VSOH HTI RUN ALC-1

Yield of wax obtained in laboratory dewaxing test with acetone at the specified (a) temperature.

SUMMARY OF 'H-NMR ANALYSES AND MICROAUTOCLAVE TESTS RELEVANT TO RUN ALC-I DEWAXING OPERATIONS

(a) Modified equilibrium tests with **Old Ben Mine** coal

Black Thunder Mine coal (4.55 g), solvent (5.45 g), **30 min, 440** "C **(824** OF), **10.3 MPa (1500 psig) H, (cold) (4**

Data shown from period 20 0-6 bottoms sample **(c)**

TABLE 14

ANALYSIS AND DE-OILING YIELDS OF WAX FROM RUN ALC-1, PERIOD 25

(a) Average of **four determinations. Repeatability was ca.** **Soh* **relative.**

(b) Suspect low, cf. **text.**

CONSOL ANALYSES OF PB-03 CRUDE OIL ASSAY FRACTIONS

BD = **Below Detection Limit**

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Temperature Equivalents:

 $70 °F = 21 °C$ **180 °F = 82 °C 350 "F** = **177 "C 400 "F** = **204 "C 650 "F** = **343 "C** $550 °F = 288 °C$

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RESID REACTIVITY TESTS

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RESID REACTIVITY TESTS

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COMPARISON OF RESID CONVERSION DATA OBTAINED IN 45 mL MICROAUTOCLAVES (CONSOL R&D) AND SHORT TIME BATCH REACTOR (UNIVERSITY OF DELAWARE)

Component methane ethane ethylene propane propylene n-butane I-butane 1 -butene trans-2-butene cis-2-butene n-pentane I-pentane carbon monoxide carbon dioxide nitrogen argon hydrogen **Volume,** % 8.0 **3.0** 0.5 2.0 0.5 1 .o 0.5 0.5 0.5 0.5 0.5 0.5 1 .o 1 .o 0.5 1 .o 78.5

GAS CHROMATOGRAPHY CALIBRATION GASES

HTI Run ALC-1 Component Distribution Of Feed Slurry Samples

^{*} Toluene Extracted

Figure 3. Component Distribution of Pressure-Filter Cake (PFC) Samples from HTl Run ALC-1.

HTI RUN ALC-1 % PARAFFINIC H -WHOLE SAMPLES

HTI RUN ALC-1 MICROAUTOCLAVE COAL CONVERSION

Figure 6. Donor Solvent Quality of Selected Whole and Distillate Samples from HTI Run ALC-1.

HTI Run ALC-1 Phenolic OH Content Of Whole Samples and Resids

Figure 7. Phenolic -OH Concentration in **Process** Stream Samples from HTI Run ALC-1.

[APPENDIX 1](#page-5-0)

RECALIBRATION OF FTlR SPECTROSCOPIC PHENOLIC -OH METHOD

RECALIBRATION OF FTIR SPECTROSCOPIC METHOD FOR PHENOLIC -OH DETERMINATION

INTRODUCTION

Replacement of CONSOL's original Fourier-transform infrared (FTIR) spectroscopic system (Nicolet 170SX) with a new model (Nicolet Magna 550) required that the FTlR method used for determination of phenolic -OH concentration in liquefaction samples be recalibrated for the new system. Implementation of the method on the new system also required that the method be reprogrammed to facilitate analysis and calculations performed on the new system. To accomplish these tasks, several steps were performed. Solutions were prepared of six phenol standards at four concentrations each. The spectrum of each standard solution was run on the new FTlR instrument system, and the spectra were visually inspected for quality. The net absorbance of the phenolic -OH peak was measured using a re-implemented version of the software. The absorbance-concentration relation of each standard was checked statistically for consistency with the other concentrations of each standard. A regression was used to obtain the new calibration relationship, and the calbration was tested by analyzing authentic coal liquids that were run on both the old and new systems.

METHOD DETAILS AND VALIDATION

The six model compounds that were previously used as standards¹⁻³ were selected as standards for recalibration, except that 2,3,5-trimethyl phenol was substituted for the 2,4,5-trimethyl phenol used in the original calibration. Four solutions of each phenol standard were prepared and run on the new FTIR instrument. The spectra were visually inspected for quality, and repeats were run as needed. The solution concentrations were randomly designed to span a range of 8.5-50.0 mm-meq/L in concentration times pathlength (PL) to cover the range of the original calibration. The solution concentration of phenolic -OH was calculated from the spectrum of each standard solution using the original calibration equation

*conc*PL (in mm-meq/L)* = *53.61 *net absorbance* - *0.1892* **(i** *1.8 as the standard error* of *the Y estimate),*

where net absorbance is the height of the phenol O-H stretch peak (found in the 3200-3400 cm⁻¹) region) minus the baseline absorbance (integration over the 3650-3750 cm⁻¹ region). From this, the sample concentration is obtained by

cone (in meq/g) = *(conc*PL) "(solution Vol* in *L)/((PL in mm) *(sample* **wt** in *g))*

Because the sample (weight) concentration of phenolic -OH in each phenol standard should be independent of solution concentration, this value (meq OH/g sample) provides a measure of agreement within the set of four solutions of each standard. These values were subjected to the Q-test for statistical outliers (at the 90% and higher confidence level). Questionable solutions of each standard were remade and rerun to obtain a new result to replace each outlying result. After a good set of data was obtained, regressions of net absorbance on conc*PL and vice versa were calculated using the four data points for each of the six standard phenols (i.e., **six** sets of regressions were calculated). Regressions of absorbance on conc*PL give the extinction coefficient or absorptivity of each of the phenols as the slope, whereas regressions of conc*PL on absorbance give the slope and intercept in a form useful as a calibration line for analysis. The

six individual regressions of four points each gave regression coefficients $(R²)$ of 0.993 to 0.999. whereas the overall regression of 24 points gave a regression coefficient **(R2)** of 0.955 (see Table 1). The individual regressions assured that the solution standards were of high quality. [Figure](#page-158-0) 1 shows all the standard data, with identification of the subsets of different phenol standards. The tabulated data are given in Table 2.

The new calibration is given as

*conc*PL (in mm-me@)* = 50.20(k2.32) **net absorbance* + *1.19* $(k2.63$ *as the standard error of the Y estimate),* $R^2 = 0.955$

from which the sample concentration is obtained as described above. The calibration error (standard error of the Y estimate) is higher than in the original calibration; this is discussed later.

In order to validate the calibration, a set of 10 samples was analyzed both on the old instrument with the old calibration, and on the new instrument with the new calibration. The sample set included several representative coal liquids, and a few model compounds (standards). For three of the samples, runs were made in duplicate on the new instrument and weighted averages were used to determine bias (i.e., each replicate contributed one-half as much to the total error as each single determination). The validation results are shown in Table 3. Figure 2 shows a parity plot of new versus reference (original) concentrations. In general, the peak position reported by the new calibration was biased 1.8 cm⁻¹ lower than the original data (range 1-3 cm⁻¹ lower).

For the entire set (3 standards, 2 high concentration samples, 5 normal concentration samples), the concentration bias was 0.00 meq/g, and the regression equation was

new conc. = *old conc.* * *1.01*(±0.02) - 0.04 (±0.19 as the standard error of the Y estimate), $R^2 = 0.997$

For the set without the standards (2 high concentration samples, 5 normal concentration samples), the concentration bias was -0.08 meq/g,and the regression equation was

new conc. = *old conc.* $*$ 0.92(\pm 0.01) $*$ 0.02 (\pm 0.08 as the standard error of the Y estimate), $R^2 = 0.998$

For the set of normal concentration samples (5 normal concentration samples), the concentration bias was 0.02 meq/g, and the regression equation was

new conc. = *old conc.* * $1.04(\pm 0.01) + 0.003 (\pm 0.01)$ as the standard error of the Y estimate), $R^2 = 0.999$

The results indicate a trivial bias of less than 0.02 meq/g in magnitude for the samples. The phenolic extract sample with a low new result created a larger magnitude bias when the standards were omitted from the validation. The discrepancy between old and new results was larger for samples with higher phenolic -OH concentrations (i.e. >I *.O* meq/g, which is outside the most common range for coal liquids). The largest concentration differences observed between the old instrument and new instrument amounted to a relative error of ca. 7%. The calibration regression showed **a** larger error than the original calibration, but the new one involved more standards and was much more rigorously set up from a statistical standpoint. It is likely that the earlier calibration underestimated the actual error. It appears that the calibration regression also represents a relative error of ca. *5%,* but this error analysis may not be rigorous. **A** calculation of the calibration error for each sample was implemented in the new version of the method, according to the following formula:

conc error due to calibration (in meq/g) = *2.63*(solution Vol in L)/((PL in mm)*(sample wt in g))*

This approach essentially disregards the error in measurement of solution volume, sample weight, and cell pathlength, and assumes that the error arising from differences in the model compounds used as standards is the largest error contributor. A printout of the computer program used to process each spectrum and perform the concentration calculations and a printout of the report output are shown in Listings 1 and 2. Because the old FTlR command set was not fully implemented on the new FTlR system, software implementation of the method on the new instrument was not straightforward. This required an advanced macro language (Visual Basic 3.0) implementation, rather than the simpler standard macro language provided by Nicolet. ation Vol in L)/((PL in mm) "(sample wf in g))
In measurement of solution volume, sample weight,
The ansing from differences in the model compounds
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check as the odd FTIR co

REFERENCES

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- 2. Burke, **F.** P.; Winschel, R. A.; Robbins, G. A. "Recycle Slurry Oil Characterization, Final Report, October 1980 through March 1985", DOE/PC 30027-61, March 1985.
- **3.** Robbins, G. A.; Winschel, R. A.; Burke, F. P. "Phenolic -OH as a Process-Performance Indicator in Two-Stage Liquefaction", Am. Chem. Soc. Div. Fuel Chem. Prepr. 1985, 30(4), 155-163.

TABLE AI -1

REGRESSION RESULTS FOR FTlR SPECTROSCOPIC PHENOLIC -OH METHOD RECALIBRATION

TABLE AI -2

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 $\mathcal{L}_{\rm{max}}$ and $\mathcal{L}_{\rm{max}}$

STANDARDS DATA FOR FTlR SPECTROSCOPIC PHENOLIC **-OH METHOD RECALIBRATION**

TABLE A1-3

VALIDATION DATA FOR FTlR SPECTROSCOPIC PHENOLIC -OH METHOD RECALIBRATION

*Peak found represents a non-phenolic component. The concentrations are reported here as 0.00 meq/g for the purpose of determining bias; a more accurate representation of the concentration is "none detected".

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RE-STANDARDIZATION OF FTlR PHENOL METHOD

Figure **AI-I.** Calibration Data Showing Subsets of Each Standard and Overall Regression Line.

PARITY PLOT FOR PHENOL METHOD VALIDATION

Figure **A1-2.** Parity Plot of Validation Results Obtained from New and Original Spectrometer and Calibration.

Listing Al-1 - Example Report Output for FTIR Spectroscopic Phenolic -OH Determination

>>>>> START OF PHENOLIC -OH DETERMINATION <<<<< TODAY'S DATE AND TIME: 4/1/97 10:32:15 AM SPECTRUM FILENAME: C:\OMNIC\SPECTRA\PHENOLS\PHEN0190.SPA SPECTRUM TITLE: HTI ALC-1 SOH 1STG OIL PER.25-WHOLE

Phenol peak height is .592 Abs. (Peak Height Method) Phenol peak location **is** 3307.373 cm-1 (Peak Height Method)

MaxPeak: Normal Completion, 1 Peak(s) Found Phenolic -OH peak height is .592 Abs. Phenolic -OH peak location is 3307.72 cm-1 Max. Peak Position is 3307.72 cm-1 Max. Peak **is** 592 Abs. Baseline is *.00055* Abs., based on integration over 3750-3650 cm-1 Net absorbance is .5914 Abs. Conc*PL for Phenolic -OH is 30.8783 mm meq/g Solution Sample Volume **is** 2.00 mL Sample Weight is 0.0463 g **IR** Cell Pathlength is 1 .OO mm Amount of Phenolic -OH is 1.334 +/- 0.1 14 milliequivalents/gram of sample (meq/g) >>>> END OF PHENOLIC -OH REPORT *<<<e<*

Listing **A1 -2** - Computer Program Listing for FTlR Spectroscopic Phenolic **-OH** Determination

Notes: Written in Visual Basic 3.0, Requires Nicolet Macros Pro and Nicolet OMNIC Software, Used on Nicolet Magna 550 Spectrometer

[PHENCLCl-FRM](#page-5-0) - **1**

VERSION 2.00 Begin Form Form1
Caption = " = "Phenolic -OH Determination" ClientHeight = 5865 $ClientLeft$ = 60
 $ClientTop$ = 1680 ClientTop $ClientWidth = 7545$
Height = 6270 $Height$ = 6270
Left = 0 Left *=o* LinkTopic = "Forml" ScaleHeight = 5865 ScaleWidth = 7545
Top = 1335 $Top = 1335$
 $Width = 7665$ $= 7665$ Begin CommandButton Command2
Caption = "Calculate Results = "Calculate Results"
= 375 Height
Left $= 240$ T ablndex [= 10](#page-14-0)
 T op = 4320 $Top = 4320$
 $Width = 1815$ $= 1815$ End Begin TextBox Text1
Height = 217 $Height$ = 2175
Left = 240 $= 240$ MultiLine = **-1** 'True [ScrollBars = 3](#page-80-0) 'Both
Tablndex = 9 Tablndex
Text $=$ "Text1" $\begin{array}{rcl} Top & = & 120 \\ Width & = & 709 \end{array}$ $= 7095$ End Begin TextBox Text8 Backcolor = &H0080FF80& Height = 285
Left = 360 $= 360$
 $= 8$ Tablndex
Text $Top = 3840$
 $Width = 7095$ $= 7095$ $=$ "Text8" End Begin TextBox Text7 $BackColor = $&H0080FFFF&Height = 285$$ $= 285$ $Left$ = 360 T ablndex = 7 Top = 3480
 $Width$ = 3495 $= 3495$ $Text = "Text7"$

End Begin TextBox Text6
Height = 285 Height = 285
Left = 5760 $= 5760$ ScrollBars = 2 'Vertical
Tablndex = 5 Tabindex
Text $Top = 3480$
 $Width = 1695$ $= 1695$ $=$ "Text6" End Begin TextBox Text5
Height = 285 Height
Left $= 3960$
 $= 4$ Tablndex
Text $=$ "Text5" **PHENCLCl .FRM** - **2** $Top = 3480$
 $Width = 1695$ $= 1695$ End Begin TextBox Text4
Height = 285 Height = 285
Left = 3240 $= 3240$
 $= 3$ Tablndex
Text $=$ "Text4" $Top = 3120$
 $Width = 3495$ $= 3495$ End Begin TextBox Text3
Height = 285 $Height$ = 285
 $Left$ = 3240 $= 3240$
= 2 Tablndex
Text $=$ "Text3" **TOP** = 2760 $= 3495$ End Begin TextBox Text2 Height = 285
Left = 3240 $= 3240$
= 1 Tablndex
Text $=$ "Text2" $Top = 2400$
 $Width = 3495$ $= 3495$ End Begin CommandButton Command1 Caption = "Exit When Done"
Height = 375 Height Left $= 4920$ Tablndex $= 0$
Top $= 432$ $Top = 4320$
 $Width = 2535$ $= 2535$ End Begin Label Label3 Alignment = 1 'Right Justify
BackColor = &H000000C0& BackColor = &H000000C0&
Caption = "Pathlength of IR Caption = "Pathlength *of* IR Cell in **mm:"**

ForeColor = &H00FFFFFF&
Height = 255 Height = 255
Left = 720 $= 720$
 $= 6$ Tabindex
Top Top = 3120
Width = 2415 $= 2415$ End Begin Label Label2 Alignment $= 1$ 'Right Justify BackColor = &H000000C0&
Caption = "Volume of Samp = "Volume of Sample Solution in mL:" ForeColor = &H00FFFFFF&
Height = 255 Height = 255
Left = 240 $= 240$
= 12 TabIndex
Top $Top = 2400$
 $Width = 2895$ $= 2895$ End Begin Label Label1 Alignment $= 1$ 'Right Justify BackColor = &H000000C0&
Caption = "Weight of Samp = "Weight of Sample in grams:" ForeColor = &H00FFFFFFF&
Height = 255 Height
Left $= 720$

PHENCLC1 .FRM - **3**

Tabindex Width Top End End $= 11$ = 2760 = 2415

PHENCLCl .FRM - **1**

Dim Results(100, 2), ResultStr, TotalLen, StartStr, EndStr, NumPeaks, MaxPeak, PeakPo Dim NetAbs As Double, BslAbs As Double Dim AmntPhen As Variant, ConcPL As Variant, ErrPhen **As** Variant Dim SpecFile, SpecTitle **S**

```
Sub GetResults 0 
  StartStr = 1TotalLen = Len(ResultStr)Error Trap for 0 peaks 
  StartStr = InStr(StartStr, ResultStr, "Position:") 
  ResultStr = Mid(ResuItStr, StartStr) 
  StartStr = 10 
  EndStr = InStr(StartStr, ResultStr, "Intensity:") 
  Results(1, 1) = VaI(Mid(ResultStr, StartStr, EndStr - StartStr)) 
  PeakPos = Results(1, 1)StartStr = EndStr 
  ResultStr = Mid(ResultStr, StartStr) 
  StartStr = 11EndStr = InStr(StartStr, ResultStr, "Position:") 
If InStr(1, ResultStr, "No peaks found") > 0 GoTo NoPeaks
```
If EndStr = 0 Then EndStr = Len(Resu1tStr) Results(l,2) = Val(Mid(ResultStr, StartStr, EndStr - StartStr)) $MaxPeak = Results(1, 2)$ StartStr = EndStr ResultStr = Mid(ResultStr, StartStr) TotalLen = Len(ResultStr) PeakNo = 1 **If** TotalLen < 20 GoTo LoopEnd PeakNo = PeakNo + 1 StartStr = 10 EndStr = InStr(StartStr, ResultStr, "Intensity:") Results(PeakN0, 1) = VaI(Mid(ResuItStr, StartStr, EndStr - StartStr)) StartStr = EndStr ResultStr = Trim(Mid(ResuItStr, StartStr)) StartStr = 11 EndStr = InStr(StartStr, ResultStr, "Position:") If EndStr = 0 Then EndStr = Len(ResultStr) Results(PeakN0, 2) = Val(Mid(ResultStr, StartStr, EndStr - StartStr)) If (Results(PeakNo,2) > MaxPeak) Then ' Stop here if only 1 peak Loopstart:

PHENCLCI .FRM - **2**

MaxPeak = Results(PeakNo,2) PeakPos = Results(PeakN0, 1) End If ResultStr = Mid(ResuItStr, EndStr) TotalLen = Len(ResultStr) If (TotalLen > 30) GoTo Loopstart Else GoTo LoopEnd NoPeaks: Text6.Text = "MaxPeak: No peaks were found" GoTo SubEnd Text6.Text = "MaxPeak: Normal Completion, " & PeakNo & " Peak(s) Found" Text1.Text = Text1.Text & Chr (13) & Chr (10) & Text6.Text Text6.Text = "Phenolic -OH peak height is " & MaxPeak & " Abs." Text1.Text = $Text1.Text$ & $Chr(13)$ & $Chr(10)$ & $Text6.Text$ Text6.Text = "Phenolic -OH peak location is " & PeakPos & " cm-1" Text1.Text = $Text1.Text$ & $Chr(13)$ & $Chr(10)$ & $Text6.Text$ ExecuteOMNlC "CorrectedPeakArea 3800 3750 3650 3600" Result\$ = GetOMNIC("Result Current") Text6.Text = "Max. Peak Position is " & PeakPos & " cm-1" Text1.Text = $Text1.Text$ & $Chr(13)$ & $Chr(10)$ & $Text6.Text$ Text6.Text = "Max. Peak is " & MaxPeak & " Abs." Text1.Text = $Text1.Text$ & $Chr(13)$ & $Chr(10)$ & $Text6.Text$ BslAbs = GetVal(Result\$, "Uncorrected:") / 100# Text6.Text = "Baseline is " & BslAbs & **I'** Abs., based on integration over 3750-365 Text1.Text = $Text1.Text$ & $Chr(13)$ & $Chr(10)$ & $Text6.Text$ NetAbs = Format(MaxPeak - BsIAbs, "#0.00##") Text7.Text = "Net absorbance is " & NetAbs & **'I** Abs." Text1.Text = $Text1.Text$ & $Chr(13)$ & $Chr(10)$ & $Text7.Text$ ConcPL = Format((50.2 * NetAbs) + 1.1 9, **"###O.OOT%Y")** Text8.Text = "Conc*PL for Phenolic -OH is " & ConcPL & " mm rneq/g" Text1.Text = $Text1.Text$ & $Chr(13)$ & $Chr(10)$ & $Text8.Text$ LoopEnd: SubEnd: 0 cm-1"

Text $2.Text = "2.0"$ Text $3.Text = "1.0"$ Text4.Text = $"1.0"$ End Sub

Sub Command1 Click ()

PHENCLCl.FRM - **3**

-
- " ExecuteOMNIC "Select First"
" ExecuteOMNIC "Set Display YAxisMode FullScale"
- " Execute OMNIC "Close Window No"" Phenol sample"
- " ErrMsgBox

End

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End Sub

Sub Command₂ Click () $SampleVol = \overline{Format}(Val(Text2.Text), "#0.00")$ SampleWt = Format(VaI(Text3.Text), *"##O.OOW)* PathLen = Format(Val(Text4.Text), "#O.OO") Text6.Text = "Solution Sample Volume is " & SampleVol & " mL" Text1.Text = $Text1.Text 8 Chr(13) 8 Chr(10) 8 Text6.Text$ Text6.Text = "Sample Weight is " & SampleWt & " g" Text1.Text = Text1.Text & Chr(13) & Chr(10) & Text6.Text Text6.Text = "IR Cell Pathlength is " & PathLen & " mm" Text1.Text = $Text1.Text$ & $Chr(13)$ & $Chr(10)$ & $Text6.Text$ AmntPhen = ConcPL * SampleVol / (PathLen * SampleWt * 1000) AmntPhen = Format(AmntPhen, "##O.OO#) ErrPhen = 2.63 * SampleVol / (PathLen * SampleWt * 1000) ErrPhen = Format(ErrPhen, "##O.OO#') Text8.Text = "Amount of Phenolic -OH is " & AmntPhen & " +/- **'I** & ErrPhen & " mill iequivalents/gram of sample (meq/g)"

Text1.Text = $Text1.Text$ & $Chr(13)$ & $Chr(10)$ & $Text8.Text$

Text6.Text = Chr(13) & Chr(10) & ">>>>> START OF PHENOLIC -OH DETERMINATION <<<<<

Result\$ = "TODAY'S DATE AND TIME: **'I** & Date & " " & Time Text6.Text = $Chr(13)$ & $Chr(10)$ & Text6.Text & $Chr(13)$ & $Chr(10)$ & Result\$ Result\$ = "SPECTRUM FILENAME: " & SpecFile Text6.Text = $Chr(13)$ & $Chr(10)$ & Text6.Text & $Chr(13)$ & $Chr(10)$ & Result\$ Result\$ = "SPECTRUM TITLE: " & SpecTitle Text6.Text = Text6.Text & Chr(13) & Chr(10) & Result\$ & Chr(13) & Chr(10) Text1.Text = $Chr(13)$ & $Chr(10)$ & Text6.Text & $Chr(13)$ & $Chr(10)$ & Text1.Text Text6.Text = **'I>>>>** END OF PHENOLIC -OH REPORT <<<<<" Text1.Text = Text1.Text & Chr (13) & Chr (10) & Text6.Text Clipboard.Clear Clipboard.SetText Text1.Text ExecuteOMNlC "StartLogging""C:\OMNIC\LOG\PHENOLS.LOG" ExecuteOMNIC "LogDisplay On" ExecuteOMNlC "Paste" **ErrMsgBox** ExecuteOMNlC "StopLogging" **ErrMsgBox** LogFile = "C:\OMNIC\LOG\PHEN96aa.txt" TotalLen = Len(Result\$) Open LogFile For Append Access Read Write Shared As #1 Len = TotalLen
Result\$ = Text1.Text Print #1, Result\$: Close #1

Text6.Text = "Last Error was " & Err End Sub

PHENCLCI .FRM - **4**

Sub Form Load 0 Load OmTalk

I% = StartOMNIC(1, "") ExecuteOMNlC "DisplayLimits 4000 400 0.0 1.2" ExecuteOMNlC "Set Display ScaleAll True" SpecFile = GetOMNIC("Spectrum Filename") SpecTitle = GetOMNIC("Spectrum Title") ExecuteOMNlC "Set Display Mode OverlayMode" ExecuteOMNlC "Set Display YAxisMode FullScale" ExecuteOMNlC "DisplayLimits 4000 3000 0.0 1.2" ExecuteOMNlC "Set Display ScaleAll True" ExecuteOMNlC "Set Display YAxisMode FullScale" ExecuteOMNlC "CopySelectedSpectra" ExecuteOMNIC "LogDisplay On" ExecuteOMNlC "Paste" ExecuteOMNlC "Smooth 25" ExecuteOMNlC "PeakHeight 3300 Shift" **ErrMsgBox** Result\$ = GetOMNIC("Result Current") PhenPeak = GetVal(Result\$, "Y:") PhenPos = GetVal(Result\$, "X:") Textl .Text = "Phenol peak height is " & PhenPeak & " Abs. (Peak Height Method)" Text1.Text = Text1.Text & Chr(13) & Chr(10) & "Phenol peak location is " & PhenPo Text1.Text = Text1.Text & Chr (13) & Chr (10) ExecuteOMNlC "Set Display YAxisMode FullScale" ExecuteOMNlC "DisplayLimits 3400 3200 0.0 1.2" ExecuteOMNlC "Set Display YAxisMode FullScale" ExecuteOMNlC "MinMax" Result\$ = GetOMNIC("Result Current") YMin = GetVal(Result\$, "Min:") YMax = GetVal(Result\$, "Max:") ExecuteOMNlC "CustomScale YMin YMax" ExecuteOMNlC "Display" ExecuteOMNlC "PeakPick 0.0 100" **ErrMsgBox** ResultStr = GetOMNIC("Result Current") $Text2.Text = "$ Text3.Text = "" Text6.Text = "" **GetResults s** *8,* " cm-1 (Peak Height Method)"

```
End Sub
```
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APPENDIX 2

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SAMPLE REQUEST FOR HTI RUN PB-05

CONSOL Inc. Research & Development 4000 Brownsville **Road** Library, PA 15129-9566 412-854-6600 **Fa:** 412-854-6613 **4 12-854-6748**

August 2, 1996

Dr. V. Pradhan Hydrocarbon Technologies, Inc. **P.** 0. Box 6047 New York and Puritan Avenues Lawrenceville, NJ 08648

Dear Vivek:

Our sample request for Run PB-05 is described below. We are requesting material in sufficient quantities to allow for sample distribution to other research groups, as needed. We understand that you cannot always provide the requested amounts, and we can work with smaller quantities.

We would like to receive from each run condition: 1) 250 g of separator overhead (SOH) oil; 2) 250 g of atmospheric still overhead (ASOH); 3) 350 g of continuous atmospheric still **(CAS)** bottoms or its equivalent, such as 0-6 bottoms; **4)** 350 g of pressure-filter liquid (PFL); *5)* 350 g of pressure-filter cake (PFC); 6) 350 **g** of feed slurry; and 7) 350 g of interstage sample (first-stage product). Please also include: *8)* 250 g of the start-up/make-up oil from the beginning of the run; **9)** 300 g sample of the feed coal; 10) 300 g sample of the feed resid; 11) 2 kg sample of the unprocessed feed auto-fluff; 12) 1 kg sample of the simulated **MSW** plastics; and 13) 250 g of SOH oil from any operating periods in which the in-line hydrotreater **is** bypassed.

For these samples, we prefer aliquots of the same samples HTI is using for material balance workups (i.e., from the last period of a run condition).

Let us know of any problem areas with this request. Thank you for your assistance.

Sincerely,

G. A. Robbins Sr. Research Chemist

/Is

cc: R. M. Statnick R. A. Winschel S. **D.** Brandes A. G. Comolli - HTI M. A. Nowak - PETC E. *8.* Klunder - PETC

APPENDIX 3

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CRUDE OIL ASSAYS OF NET PRODUCTS OF HTI RUN PB-03

(Periods 6, 7, and 8 = Condition 2) (Periods 9, 10, and 11 = Condition 3)

CONSOL lnc. Research & Development 4000 Brownsville Road Library, PA 15129-9566 **4** 12-854-66OO FAX: 412-854-6613

4 1 2-854-6683

November 15, 1996

Dr. M. A. Nowak U.S. Department **of** Energy Pittsburgh Energy Technology Center P.O. Box 10940 Pittsburgh, PA 15236

Subject: DOE Contract DE-AC22-94PC93054

Dear Mike:

Enclosed are the reports on the crude oil assays of the two products of HTI Run PB-03. The online hydrotreater was in use during Periods 6, 7, and 8 (sample HTI PB-03-6, -7, and -8), and it was by-passed during periods 9, 10, and 11 (sample HTI PB-03-9, -10, and -11). CONSOL's preparation of the samples is described in the August 1996 status report for the subject contract.* These reports will be included in one of our Technical Progress Reports.

Sincerely,

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*and in **the** main bvdy of this report.

R. A. Winschel Research Group Leader Exploratory Research Group

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cc: **A.** G. Comolli - HTI P.-2. Zhou - BRSC

11/04/96

Consol, Inc. 4000 Brownsville Road Library, PA 15129-9566

Attention: R.A. Winschell

Reference: Crude Assav on "HTI PB-03-6,7,8" received in September 1996

Mr. Winschell,

The following results are based on a composite of (2) separate distillations performed on this sample. The Distillation Data Report provides a breakdown of the separate distillations. Should you have any questions concerning this report, you may contact me at **(713) 844-331 1,** or by fax at **(713) 844-3330.**

Sincerely,

Robert Kelly Distillation Manager

November 12, 1996

Houston, Texas

Our Reference : H0/96-005162 Your Reference : PO# 01 -001 -033668

Consol, Inc. Research & Development 4000 Brownsville Road Library, PA 151 29-9566

ATTN: R.A. Winschel

Reference: To perform "CRUDE ASSAY" on "HT1 PB-03-6,7,8" received in September 1996.

Dear Mr. Winschel:

Please find enclosed the original report on the above referenced submitted sample(s) and our invoice for services rendered.

Should you have any questions regarding this report, please do not hesitate to contact us at your convenience.

We trust you find all in order and thank you for requesting our services.

Very Truly Yours,

INCHCAPE TESTING SERVICES CALEB BRETT U.S.A.

John Mowrey

Laboratory Manager

Enclosures: JM/rm

9809 Rowlett Road Houston, TX 77075 Phone: (713) 946-2420 Fax: (713) 946-0545

Your Ref: PO# 01-001-033668 Date: 8-NOV-1996

[Laboratory Report No. 96-005162-0-HOUS; 1](#page-5-0)

Consol, Inc. **4000 Brownsville Road Library, PA 15129-9566**

For the Attention of R. A. Winschel

(TOTAL NUMBER OF PAGES 8)

Approved by: John Mowrey

N-Heptane Insolubles

Boiling Point Distribution

wt. %

D3279 D5307 **0.05**

Attached

e
 e- **Inchcape Testing Services**
 E- Caleb Brett **Caleb Brett Caleb Brett Caleb 2**

Laboratory Report No. 96-005162-0-HOUS; 1 - [Page 3](#page-80-0) of 8

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65.7 0.7175 84.48 15.51 **co.01** 35.5 8.4

$\begin{array}{c} \begin{array}{c} \begin{array}{c} \end{array} \\ \end{array}$ **__** Inchcape Testing Services

Caleb Brett

Laboratory Report No. $96-005162-0-HOUS; 1 - Page 4 of 8$ $96-005162-0-HOUS; 1 - Page 4 of 8$ $96-005162-0-HOUS; 1 - Page 4 of 8$

$\begin{aligned} \n\end{aligned}$ **Inchcape Testing Services**

Laboratory Report No. 96-005162-0-HOUS; 1 - [Page](#page-71-0) *5* of 8

Laboratory Report No. $96-005162-0-HOUS; 1 - Page 6 of 8$ $96-005162-0-HOUS; 1 - Page 6 of 8$ $96-005162-0-HOUS; 1 - Page 6 of 8$ $96-005162-0-HOUS; 1 - Page 6 of 8$

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Inchcape Testing Services

Laboratory Report No. $96-005162-0-HOUS; 1 - Page 7 of 8$ $96-005162-0-HOUS; 1 - Page 7 of 8$ $96-005162-0-HOUS; 1 - Page 7 of 8$

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[Laboratory Report No. 96-005162-0-HOUS; 1](#page-5-0) - **[Page 8 of 8](#page-12-0)**

[7](#page-11-0)

Boiling Point Distribution
ASTM D-2887

_.. --____- -- - ___ __ - - - . _.~____

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Start Time: 0.2 minutes End Time: 24.3 minutes Area: 139559232.0 Slice Width: 0.80 sec

 μ [Sample Offset: 21](#page-25-0)859.9 Baseline Offset: 22729.7 Calibration File: 1014rt Calibration Date: 10/16/96

Baseline Subtracted: c:\tc4\gc10\1014b

-

أن المساوية المساوية
1. FILE: c:\tc4\sd6890\5162-7 NOV 11, 1996 - 17:24:24

' FILE: ~:Uc4\sd6890W162-7 **SlMDlS** EXPERT V5.0 I PARAMETER: 2887 **Page** ³

Boiling Point Distribution ASTM D-2807

Start Time: 0.7 **minutes End Time:** 16.4 **minutes Area:** 329832608.0 *Slice* **Width: 0.80 sec Baseline Subtracted:** c:\tc4\sd6890\1111 **b Sample Offset:** 8544.0 **Baseline Offset:** 8792.0 **[Calibration File:](#page-5-0) 1111rt Calibration Date:** 1111 1/96

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WinAssay '95

Version 1.00

- P

Final Repotts

Distillation Summary Report

1 1

Prepared For: Sample ID:

Date:

Consol Inc. HTI PB-03-6,7,8 (Charge #1) 9/25/96

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Cut Temp Degrees F **DUMP** Specific LIQ **CUM. LIQ** $WT%$ $\overline{\text{CUM}}$ **API** MID $\overline{\bf ro}$ $WT(g)$ Gravity **MLS** VOL% VOL% WT% **GRAVITY** LIQ VOL% **ASTM D2892 Distillation Yields IBP** $\overline{70}$ 61.80 0.5780 106.92 0.94 0.94 0.65 0.65 113.31 0.47 $\overline{70}$ 180 896.90 0.7166 1251.60 11.00 11.94 9.45 10.10 65.96 6.44 180 $\overline{350}$ 3092.80 0.7803 3963.60 34.85 46.79 32.58 42.67 49.84 29.37 $\overline{350}$ 400 829.60 0.8430 984.10 8.65 55.44 $8,74$ 51.41 36.35 51.12 400 550 2613.00 0.8892 2938.60 25.84 27.52 78.94 68.36 81.28 27.63 550 $\overline{650}$ 1399.90 0.9282 1508.19 13.26 94.54 14.75 93.68 20.95 87.91 $650+$ 600.00 0.9248 648.79 5.70 100.24 6.32 100.00 21.51

Loss (Grams): 9.8 (0.10 Wt.%) Distribution: (2/3) 6.5 g to IBP-70 F (ID) 3.3 g to 70-180 F

WinAssay '95 Quality Control Applications

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Cum. Mid Vol% v. API **Gravity**

HTI PB-03-6,7,8 (Charge #I)

WinAssay '95 True Boiling Point Curve Vaporline Temperature v. Cumulative Wt% Yield

Sample ID

HTI PB-03-6,7,8 (Charge #1)

WinAssay '95 True Boiling Point Curve vs Cumulative Vol% Yield

Sample ID

ITS- Caleb Brett Distillation Group

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ASTM D2892/05236 CHARGE INFORMATION

Client Name: Sample ID: Date:

Operator: Robert Kelly

Charge Mass D2892(g): Charge *S.G* **D2892 (60/60F):**

Charge Mass D5236(g): Charge S.G. D5236 (60/60F):

Water Weight Removed (g): Initial Vapor Temp: Whole Crude Sulfur Wt%:

CALEB BRETT HOUSTON

Analyzed: 9/20/96 S:03 PM Normalized to 100.00% Reported: 09-23-1996 11:06:08

Components Listed in Chromatographic Order

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 ~ 10

A3-21

CALEB BRETT HOUSTON

Analyzed: 9/20/96 5:03 PM Normalized to 100.00% Reported: 09-23-1996 11:06:08

Composite Report Totals by Group Type & Carbon Number (in Weight Percent)

Grand Total: 100.000

Molecular Weight and Relative Density Data

File: 5174A2.DHA

CALEU BRETT **HdUSTON**

Analyzed: 9/20/96 **5:03** PM Normalized to 100.00% Reported: 09-23-1996 **11:06:08**

Composite Report Totals by Group Type & Carbon Number (in Volume Percent)

File: 5174A2.DHA

CALEB BRETT HOUSTON

Analyzed: **9/20/96 5:03** PM Normalized to 100.00% Reported: **09-23-1996 11:06:08**

Boiling Point Distribution Data

Research Octane Number **=107.33** (Calculated from Individual Component Values)

> Contribution to Total by: Paraffins: **88-70** Iso-paraffins: **17.97** Aromatics: 0.00 Naphthenes: 0.00 Olefins : 0.61 Oxygenates: *0.00*

File: **S174A2** .DHA

WinAssay '95

Version 1.00

Final Reports

Loss (Grams): 17.8 (0.20 Wt.%) Distribution: (2/3) 11.9 g to IBP-70 F (In) *5.9* **g to 70-180 ^F**

WinAssay '95 Quality Control Applications

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Cum. Mid Vol% v. API Gravity

HTI PB-03-6,7,8 (Charge #2)

WinAssay '95 True Boiling Point Curve Vaporline Temperature v. Cumulative Wt% Yield

Sample ID

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WinAssay '95 True Boiling Point Curve vs Cumulative Vol% Yield

Sample ID

HTI PB-03-6,7,8 (Charge #2)

ASTM D2892/D5236 CHARGE INFORMATION

Sample ID: Date:

Operator: Robert Kelly

Charge Mass D2892(@: Charge S.G D2892 (60/60F):

Charge Mass D5236(g):

 \sim

Water Weight Removed (g): Initial Vapor Temp: Whole Crude Sulfur Wt%:

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CALEB BRETT HOUSTON

Components Listed in Chromatographic Order

File: 516482.DHA Sample: 96-5164-2b

 $p. 1$

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CALEH BRETT HdUSTON

Analyzed: 9/20/96 6:06 PM Normalized to 100.00% Reported: 09-23-1996 11:09:47

Composite Report Totals by Group Type & Carbon Number (in Weight Percent)

Molecular Weight and Relative Density Data

File: 5164B2.DHA

CALEB BRETT HOUSTON

- 40

Analyzed: 9/20/96 6:06 PM Normalized to 100.00% Reported: 09-23-1996 11:09:47

Composite Report Totals **by** Group Type & Carbon Number (in Volume Percent)

File: 5164B2 .DHA
CALEB BRETT HOUSTON

Analyzed: 9/20/96 *6:06* PM Normalized to 100.00% Reported: 09-23-1996 11:09:47

Boiling Point Distribution Data

Research Octane Number =105.62 (Calculated from Individual Component Values)

File: 5164B2.DHA

NOV 1 1 1996 R.A.W.

11/04/96

Consol, Inc. 4000 Brownsville Road Library, PA 15129-9566

Attention: R.A. Winschell

Reference: Crude Assay on "HTI PB-03-9, 10, 11" received in September 1996

Mr. Winschell,

The following results are based on a composite of (2) separate distillations performed on this sample. The Distillation Data Report provides a breakdown of the separate distillations. Should you have any questions concerning this report, you may contact me at **(713) 844-3311,** or **by** fax at **(713) 844-3330.**

Sincerely,

Robert Kel Distillation Manager

November 12, 1996

Houston, Texas

Our Reference : H0/96-005170 Your Reference : PO# 01 -001 -033668

Consol, Inc. Research & Development 4000 Brownsville Road Library, PA 151 29-9566

ATTN: R.A. Winschel

Reference: To perform "CRUDE ASSAY" on "HT1 PB-03-9,10,11" received in September 1996.

Dear Mr. Winschel:

Please find enclosed the original report on the above referenced submitted sample(s) and our invoice for services rendered.

Should you have any questions regarding this report, please do not hesitate to contact us at your convenience.

We trust you find all in order and thank you for requesting our services.

Very Truly **Yours,**

INCHCAPE TESTING SERVICES **CALEB BRETT U.S.A.**

6John Mowrey

Laboratory Manager

Enclosures: JM/rm

9809 Rowlett Road Houston, TX 77075 Phone: (713) 946-2420 Fax: (713) 946-0545

Your Ref: PO# 01-001-033668 Date: 8-NOV-1996

Laboratory Report No. 96-005170-0-HOUS; 1

Consol, Inc. 4000 Brownsville Road Library, PA 15129-9566

For the Attention of R.A. Winschel

Approved by: John Mowrey

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Laboratory Report No. 96-005170-0-HOUS; 1 - **[Page](#page-6-0)** ² **of** *8*

Sample ID **Description**

HTI PB-03-9, 10, 11 CRUDE **OIL**

96-005170-0-HOUS-002-00 **IBP-70** Deg. F

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Laboratory Report No. 96-005170-0-HOUS; 1 - **[Page 3](#page-80-0) of 8**

1 f $\begin{aligned} \begin{aligned} \text{In theorem } \end{aligned} \end{aligned}$ The Caleb Brett **Caleb Brett**

Laboratory Report **No.** 96-005170-0-HOW; 1 - [Page 4](#page-8-0) of 8

$\frac{P}{\text{P}}$ **Pesting Services** Caleb Brett

Laboratory Report No. $96-005170-0-HOUS; 1 - Page 5 of 8$ $96-005170-0-HOUS; 1 - Page 5 of 8$ $96-005170-0-HOUS; 1 - Page 5 of 8$

$\begin{array}{c}\n\hline\n\text{Inchcap}\n\hline\n\text{Calcb Brett}\n\end{array}$ **Inchcape Testing Services**

Viscosity @ 40 C Viscosity @ 100 C Freezing Point Pour Point Aniline Point Flash Point, TCC Luminometer Number

Corrosion 3 hrs @ 122 F

Initial Boiling Point @ **5%** Evaporated @ 10% Evaporated @ 20% Evaporated @ 30% Evaporated @ 40% Evaporated @ 50% Evaporated @ 60% Evaporated @ 70% Evaporated @ 80% Evaporated @ 90% Evaporated @ 95% Evaporated Final Boiling Point

Net Heat of Combustion Total Acid Number

Smoke Point Paraffins Olefins Naphthenes Aromatics Total N & *A* Naphthalenes

Existent Gum Tube Rating Pressure Drop Cetane Number

Recovery Residue Loss

Laboratory Report No. 96-005170-0-HOUS; 1 - [Page 6](#page-10-0) of 8

cst cst Deg. F Deg. F Deg. F Deg. F

D445 D445 D2386 D97 D611 D56 D1740 D1322 G.C. G.C. G.C. G.C. D1840 D130 D381 D3241 D3241 D613 D86

006-00

16.8 0.9540 83.85 10.15 0.12 8159.0 5972 126 73.47 4.216 1.199 -19.0 -33.0 18.5 >200 22 9 13.95 4.40 20.97 60.68 81.65 14.92 la 40 4.0 <1.0 $_{18.3}$ </sub> 412 43 3 439 444 450 457 462 471 480 484 504 514 532 99.0 0.5 0.5 18190 0.74
کم*مبرگ*

G.C.

D1405 D974

mm Vol. % VOl. % VOl. % VOl. % VOl. % VOl. % mg/100mL

 mm Hq

Deg. F Deg. F Deg. F Deg. F Deg. F Deg. F Deg. **F** Deg. F Deg. F Deg. F Deg. F Deg. F Deg. F VOl. % VOl. % VOl. % BTU/ lb mgKOH/9

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External Proper Testing Services

Laboratory Report No. $96-005170-0-HOUS; 1 - Page 7 of 8$ $96-005170-0-HOUS; 1 - Page 7 of 8$ $96-005170-0-HOUS; 1 - Page 7 of 8$

A3-43

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- **P The Services** Caleb Brett **Caleb Brett Caleb Bret**

Laboratory Report No. $96-005170-0-HOUS; 1 - Page 8 of 8$ $96-005170-0-HOUS; 1 - Page 8 of 8$ $96-005170-0-HOUS; 1 - Page 8 of 8$

SAMPLE: 96-51 70-7 **FILE:** c:Uc4\sd6890\5170-7 **PARAMETER:** 2887

 \mathbb{R}^2

Boiling Point Distribution ASTM D-2887

Start Time: 0.7 minutes **End** Time: 17.8 minutes **Area:** 272058272.0 **Slice Width: 0.80** sec **Baseline** Subtracted: c:\tc4\sd6890\1111 **b Sample** Offset: 8638.0 **Baseline** Offset: 8792.0 **[Calibration File:](#page-5-0)** [1 11](#page-15-0) [1](#page-5-0) rt **Calibration** Date: 11/11/96

A3-45

SAMPLE: 96-5170-1 **[OCT](#page-18-0)** 14, 1996 - 10:58:41

Boiling Point Distribution
ASTM D-2887

Start Time: 0.2 **minutes End Time:** 26.0 **minutes Area:** 39163064.0 **Slice Width:** 0.80 **sec Baseline Subtracted:** C:\TC4\SD6890\10138 **Sample Offset:** 7366.0 **Baseline Offset:** 7234.0 **Calibration File:** 101 3rt **Calibration Date:** 10/13/96

WinAssay '95

Version 1.00

Final Reports

 $\label{eq:1} \frac{1}{\sqrt{2\pi}}\int_{0}^{\infty}\frac{1}{\sqrt{2\pi}}\left(\frac{1}{\sqrt{2\pi}}\right)^{2}d\theta.$

 $\mathcal{A}^{\text{max}}_{\text{max}}$

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Loss (Grams): 9.2 (0.11 Wt.%) Distribution: (2/3) 6.1 g to IBP-70 F (ID) 3.1 g to 70-180 F

WinAssay '95 Quality Control Applications

Cum. Mid Vol% v. API Gravity

HTI PB-03-9,10,11 (Charge #I)

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WinAssay '95 True Boiling Point Curve Vaporline Temperature v. Cumulative Wt% Yield

Sample ID

HTI PB-03-9,10,11 (Charge #I)

WinAssay '95 True Boiling Point Curve vs Cumulative Vol% Yield

Sample ID

HTI PB-03-9,10,11 (Charge #1)

ASTM 02892/05236 CHARGE INFORMATION

Lab ID: Client Name: Sample ID: Date:

Operator: Robert Kelly

Charge Mass D2892(g): Charge S.G D2892 (60/60F):

Charge Mass D5236(g): Charge *S.G.* **D5236 (60/60F):**

Water Weight Removed (g): Initial Vapor Temp: Whole Crude Sulfur Wt%:

CALEB BRETT HOUSTON د.

Analyzed: **10/2/96 8:25** *AM* rmalized to 100.00% Reported: **10-03-1996 13:00:05**

Composite Report Totals by Group Type & Carbon Number (in Weight Percent)

Molecular Weight and Relative Density Data

File: **5170Sl.DHA. A3-53**

CALEB BRETT HOUSTON **Contract Contract Contract Contract Contract**

Analyzed: 10/2/96 8:25 *AM* **Reported:** 10-03-1996 13:00:05 **Normalized** to 100.00%

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Composite Report Totals by Group Type & **Carbon Number (in Volume Percent)**

CALEB BRETT HOUSTON

CLD: CONSOLINC Analyzed: **10/2/96 8:25** *AM* Reported: 10-03-1996 13:00:05
Normalized to 100.00%

Boiling Point Distribution Data

Research Octane Number **=106.27** (Calculated from Individual Component Values)

File: **5170Sl.DHA**

 \sim

CALEB BRETT HOUSTON TID: 96-005170-0-HOUS-002-00 CID: CONSOLINC SID: HTI PB-03-9,10,11 CRUDE OIL/IBP-70 F NID: 51870 Date: 16-SEP-1996

Analyzed: 10/2/96 8:25 **AI** Normalized to 100.00% Reported: 10-03-1996 13:00:05

Components Listed in Chromatographic Order

File: 5170S1.DHA Sample: 96-5170-s1 p. 1

Components Listed in Chromatographic Order

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File: 5170Sl.DHA. Sample: 96-5170-s1

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WinAssay '95

Version 1.00

Final Reports

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Alikuwa katika sa kalimanan sa kalimanan sa katika sa katika sa katika sa katika sa katika sa katika sa katika

Loss (Grams): 16.2 (0.21 Wt.%) Distribution: *(2D)* **10.8 g** *to* **IBP-70 F (ID) 5.4 g to 70-180 F**

Cum. Mid Vol% v. API Gravity

HTI PB-03-9,10,11 (Charge #2)

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WinAssay '95 True Boiling Point Curve Vaporline Temperature v. Cumulative Wt% Yield

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Sample ID

HTI PB-03-9,10,11 (Charge #2)

WinAssay '95 True Boiling Point Curve vs Cumulative Vol% Yield

Sample ID

HTI PB-03-9,10,11 (Charge #2)

ITS- Caleb Brett Distillation Group

 $\sim 10^{-11}$

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ASTM 02892/D5236 CHARGE INFORMATION

Lab **ID**: **Client Name: Sample ID: Date:**

Charge Mass D2892(g): Charge S.G D2892 (60/60F):

Chilrge *Miiss* **DS236(g): Charge S.G. DS236 (60/60F):**

 $\label{eq:2.1} \frac{1}{2} \int_{\mathbb{R}^3} \left| \frac{1}{2} \left(\frac{1}{2} \right) \right| \, d\mu = \frac{1}{2} \int_{\mathbb{R}^3} \left| \frac{1}{2} \left(\frac{1}{2} \right) \right| \, d\mu = \frac{1}{2} \int_{\mathbb{R}^3} \left| \frac{1}{2} \left(\frac{1}{2} \right) \right| \, d\mu = \frac{1}{2} \int_{\mathbb{R}^3} \left| \frac{1}{2} \left(\frac{1}{2} \right) \right| \, d\mu = \frac{1}{2}$

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Water Weight Removed (g): Initial Vapor Temp: Whole Crude Sulfur Wt%:

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CALEB BRETT HOUSTON

Analyzed: 10/1/96 3:56 PM Normalized to 100.00% Reported: 10-03-1996 13:02:33

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Grand Total: 100.000

Totals by Group Type & Carbon Number Composite Report in Weight Percent)

Molecular Weight and Relative Density Data

File: 5170. DHA **A3-64**

CALEB BRETT HOUSTON \mathbf{A} and \mathbf{A} and \mathbf{A} and \mathbf{A}

Analyzed: 10/1/96 3:56 PM Normalized to 100.00% Reported: 10-03-1996 13:02:33

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Composite Report Totals **by** Group Type & Carbon Number (in Volume Percent)

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UALEB BRETT HOUSTON

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CLD: **CUNSOLINC** Analyzed: **10/1/96 3:56** PM SLD: **HTI** PB-03-9,l(j, **11** CHUDE Reported: **10-03-1996 13:02:33** Normalized to 100.00%

Boiling Point Distribution Data

Research Octane Number **=108.79** (Calculated from Individual Component Values)

File: **5170** .DHA

CALEB BRETT HOUSTON

Analyzed: 10/1/96 3:56 PM Normalized to 100.00% Reported: 10-03-1996 13:02:33

Components Listed in Chromatographic Order

File: 5170 .DHA Sample: 96-5170

 $p. 1$

Components Listed in Chromatographic Order

File: 5170.DHA . **Sample: 96-5170**

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p. 2

APPENDIX 4

UNIVERSITY OF DELAWARE QUARTERLY REPORT
THE KINETICS OF COAL LIQUEFACTION DISTILLATION RESID CONVERSION

QUARTERLY REPORT 7/16/96-10/15/96

Michael T. Klein Principal Investigator

William H. Calkins Co-Principal Investigator

> He Huang Research Associate

> > and

Shaojie Wang Visiting Scientist

Center For Catalytic Science and Technology Department of Chemical Engineering University of Delaware Newark, Delaware 19716

 \hat{L}

Date Published October 29,1996

Subcontract from **CONSOL** under DOE Contract DE-AC22-94PC9305

EXECUTIVE SUMMARY

Hydroconversion runs on 15 resids supplied by CONSOL, Inc. have been completed using *3* to **1** tetralin to resid **wt.** ratio with and without *3* **wt%** sulfided molybdenum naphthenate catalyst at 420 **"C** for 30 minutes. Two of the resids from the Princeton laboratory and two catalyzed Wilsonville resids have still to be analyzed. The other resids showed 30.2 to 44.9 % conversion to material boiling below 850°F in *30* minutes in the catalyzed experiments. This can be compared to uncatalyzed hydroconversion of the same materials of 15 - 23 wt % under the same conditions. The difference in these resids which *can* account for the variation of 30 up to **45** % has not yet been investigated. Previous work using a different work up procedure showed that 0.9 **wt** % sulfided molybdenum naphthenate at the same times and temperature showed only a slight increase in conversion over the uncatalyzed system.

TGA tests run on solubilized and unsolubilized resid from both the catalyzed and thermal hydroconversion experiments show that there is little of no conversion of the unsolubilized residue to material boiling below 454 °C (850 °F). However, there is substantial conversion of the solubilized resid to lower boiling material.

The molecular modeling work is proceeding satisfactorily and the algorithm for the resid structure is in place. Necessary analytical work on the **VPO** molecular weights on the SARA fractions **is** almost complete.

QUARTERLY REPORT

HYDROCONVERSION OF COAL DERIVED RESIDS

Hydroconversion of resids

Hydroconversion experiments have now been completed on all 15 resids supplied by CONSOL Inc. under the conditions agreed upon with CONSOL on May 17, 1996. Two of the resids from the Princeton laboratory and two catalyzed Wilsonville resids have still to be analyzed. The hydroconversion conditions were 30 minute runs at 420°C in 3 to 1 tetralin to resid **wt** ratio and 1500 psig hydrogen with and without using sulfided molybdenum naphthenate catalyst (as **3** wt% molybdenum based on resid charged). Each resid was also run at ambient temperature for comparative purposes.

Hydroconversion experiments using Shell Ni/Mo catalyst are still indeterminate pending the finding of a suitable method of conversion determination. Experiments have shown that the Shell catalyst itself changes under hydroprocessing conditions, making the use of ash tracer as a way of following conversion unreliable. We are considering alternative approaches to determining conversion when using that catalyst.

Work up Procedure for Hydroprocessing Experiments

The work-up of the runs is an extensive process. The reaction mixture is filtered and the filter cake is washed with cold methylene chloride into the filtrate. The methylene chloride is then distilled out of the filtrate at low temperature (below 42°C). The low concentration of the tetralin in the methylene chloride distillate is determined by gas chromatography. methylene chloride remaining and the tetralin in the resid filtrate are also determined by gas chromatography. SimDis TGA is performed on both the filtrate and the filter cake to allow the determination of the resid conversion to product boiling below 850 \degree F. The

Resid Conversion Kinetics

Since the conversion is determined as the product of the amount of resid solubilized and the percentage of that solubilized product boiling below **850°F,** the conversion calculation is based on Equation:

$$
Conversion(<850F) = TSF \times (1 - \frac{850F +}{RSF})
$$

where the Tetralin Soluble Fraction (TSF) of resid (dry-ash-free basis) is determined by ash content in the solid residue after resid hydroprocessing; the Resid Soluble Fraction (RSF) in tetralin and the **850°F'** fraction of the **RSF** are determined by running *GC* and SimDis TGA on the liquid products, respectively.

[Table 1](#page-151-0) shows the thermal and catalytic reactor runs made and the conditions used. Table **2** shows the solubilization of the thermal and catalyzed runs as well as the room temperature controls. Table **3** shows the increase in solubilization over the room temperature controls in the thermal and catalytic runs. There is a substantial solubility of all the resids in tetralin even at room temperature. The increases in solubilization over the room temperature experiments are only of the order of 10 to 20 %. The presence of a catalyst increases the solubilization only a small amount (1 to **4%).** Tables 4 and 5 show the conversions of resid to material boiling below **850°F** for the thermal and catalyzed runs, respectively.

TGA on the solid residues using the ramp method shows that all the insoluble portion of the resids or reactor solid residues shows no material boiling below **850°F.** The hydroconversion **of** the resid apparently occurs only on the solublized resid. This conversion is only about 15 - 20 % for the thermal hydroconversion to **850°F** with the tetralin donor solvent alone. On the other hand, in the presence of **3%** molybdenum naphthenate catalyst, the conversion increases to **30.2** to **44.9** % depending upon the resid. All of the 13 Wilsonville resids and two HTI resids have been run under thermal and catalytic conditions and have been analyzed. The recovery in processing the two very sticky HTI resids was relatively poor. Two more of the Wilsonville resids converted in the presence of catalyst and two HTI resids have not yet been completely analyzed.

Several difficulties have been experienced with the *gas* chromatograph, which is used in the analytical scheme. The latest problem relates to the drifting response factor of the FID detector. We are trying to avoid this problem by using **an** internal standard, and all the samples of filtrate and distillate are now being analyzed in this way. We believe, however, that the experimental errors in the conversion data presented in this report cannot be more than *5%* (relative). With the use of internal standards, we expect to minimize them.

MODELING OF RESID EYDROCONVERSION

Analysis of SARA Samples

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To date molecular weights have been determined for each of the aromatic, aliphatic and preasphaltene fractions of the SARA separations (Table *6).* However, the resin samples which were dissolved in dichlorobenzene yielded results which are too **high** and these samples are unrecoverable. It will be assumed that these samples have a *MW* of 500. The value of 500 falls within 2 experimental standard deviations of every measured resin, so this assumption should not lead to much error. To account for this approximation, the constraint in the structural model for these resids which requires the molecular weight of the resin fraction to match the experimentally determined value will be relaxed.

The molecular weights for the first **six** saturate fractions have also been determined. Only the remaining nine saturate fractions need to be determined. All other analytical information has been supplied by CONSOL, Inc.

Modeling

A molecular structure model for coal resid has been assembled (see [Table 7](#page-157-0) and Figure 1). Currently, a molecular representation of the first resid *(W258* **V-13 1B)** is being optimized to **the** analytical properties of the initial feedstock. The molecules have been defined by their basic structural attributes (e.g. number of aromatic rings, number of thiophenic rings, number of sidechains, number of unit sheets, etc.) and a construction algorithm is presented in [Figure](#page-158-0) **1.** The attributes are represented by probability distribution finctions (pdf **s)** which give the probability of finding a value or less of a given attribute. [Table 7](#page-157-0) lists the distributions used in this algorithm.

By stochastic (Monte Carlo) sampling of the pdf **s,** a molecular sample of a resid can be constructed. The properties of this representation can be compared to experimental values. The parameters of the pdf's are optimized to minimize the difference between the predicted and experimental values. The objective function used to optimize these parameters is:

$$
\chi^2 = \left(\frac{Hwt_{\text{exp}} - Hwt_{\text{pred}}}{0.0016}\right)^2 + \left(\frac{Cwt_{\text{exp}} - Cwt_{\text{pred}}}{0.0064}\right)^2 + \left(\frac{Owt_{\text{exp}} - Owt_{\text{pred}}}{0.0150}\right)^2 + \left(\frac{Mwt_{\text{exp}} - Nwt_{\text{pred}}}{0.0011}\right)^2 + \left(\frac{Swt_{\text{exp}} - Swt_{\text{pred}}}{0.0010}\right)^2 + \left(\frac{Marm_{\text{exp}} - Marm_{\text{pred}}}{0.02}\right)^2 + \left(\frac{Sarm_{\text{exp}} - Sarakt_{\text{pred}}}{0.03}\right)^2 + \left(\frac{Harn_{\text{exp}} - Harn_{\text{pred}}}{0.02}\right)^2 + \left(\frac{Harm_{\text{exp}} - Harn_{\text{pred}}}{0.02}\right)^2 + \left(\frac{Hcan_{\text{exp}} - Hcan_{\text{pred}}}{0.02}\right)^2 + \left(\frac{Hann_{\text{exp}} - Hann_{\text{pred}}}{0.02}\right)^2 + \left(\frac{Phox_{\text{exp}} - Phox_{\text{pred}}}{0.02}\right)^2 + \left(\frac{Hann_{\text{exp}} - Hann_{\text{pred}}}{0.02}\right)^2 + \left(\frac{Phox_{\text{exp}} - Phox_{\text{pred}}}{0.02}\right)^2 + \left(\frac{Phox_{
$$

The numerator of each term represents the residual between the measured and predicted values. The denominator is a weighting factor equal to approximately one standard deviation. The weighting factor ensures that values which are measured with higher accuracy are represented more closely than values measure with lower accuracy.

The model will be fine-tuned with the first few resids. In particular, an investigation into the importance of conditional probability and into the approximation of the intersheet linkages will be made. It may be necessary to make some slight alterations to the current model to get an optimal representation. Later, some reaction product data may also be added to the objective function.

Sample molecules from this representation will be examined with a view to selecting model structures for determining hydroconversion rate constants in the SCTBR reactor. Such data is not essential to the functioning of the model. However, it could provide a useful starting point for the reaction rate constants.

SUMMARY AND CONCLUSIONS

1. **A** distinct increase in the production of material boiling below **850°F** is observed in the presence of **3 wt%** sulfided molybdenum naphthenate catalyst under the conditions used. This is substantially higher than that observed in hydroprocessing in the absence of catalyst (thermal) or when only 0.9 **wt%** of the same catalyst is used.

2. occur even in the presence of the catalyst. Little or no conversion of the unsolubilized resid to lower boiling material appears to

3. The solubilization of the resid in tetralin is substantial even at room temperature. It increases only 10 to 20 wt% under hydroprocessing conditions without catalyst. There is very little increase in solubilization in the presence of catalyst. However, the breakdown of the soluble resid into low boiling material depends strongly on the catalyst.

4. These values are being correlated with analytical data of the resids. Conversions of various resids under catalytic conditions have varied from 30 to **45** wt%.

5. optimized. **A** molecular structure model for coal derived resid has been assembled and is being

6. Molecular weight determination of the various **SARA** fractions is almost complete.

FUTURE **PROGRAM**

1. In consultation with CONSOL Inc., one or more selected resids will be hydroprocessed over a range of conditions, i.e., time, temperature, tetralin to resid ratio, catalyst concentration, hydrogen pressure and solvent type (including Wilsonville recycle solvent) to establish what conversion of recycle resid can be expected under actual process conditions.

2. An effort will be made to correlate conversion data with the available structural information (NMR, elemental analysis etc.).

3. If time permits, an effort will be made to relate hydroprocessing performance in the presence of molybdenum naphthenate catalyst to that of the use of sulfided Ni-Mo/Al₂O₃ catalysts used at Wilsonville and the other promising catalysts developed in recent years.

4. hydroprocessing to assist in the molecular modeling work. If time permits, some rate constants of selected model compounds will be determined for

PUBLICATIONS

Three papers were presented to the New Orleans National Meeting of the American Chemical Society in the Spring of 1996 that relate in part to work done under this project. They are:

- **1.** He Huang, Keyu Wang, Shaojie Wang, M.T. Klein, and W.H. Calkins 'Applications of the Thermogravimetric Analysis in the Study of Fossil Fuels', *Prepr. Pap.* - *Am. Chem. Soc., Div. Fuel Chem.* **1996**, 41(1), 1.
- *2.* Keyu Wang, Shaojie Wang, He Huang, M.T. Klein, and W.H. Calkins 'A Novel Smoothing Routine for the Data Processing in Thermogravimetric Analysis', *Prepr. Pap.* - *Am. Chem. Soc., Div. Fuel Chem.* **1996**, 41(1), 27.
- **3.** He Huang, Keyu Wang, Shaojie Wang, M.T. Klein, and W.H. Calkins 'Distillation of Liquid Fuels by Thennogravimetry', *Prepr. Pup.* - *Am. Chem. Soc., Div. Fuel Chem. 1996,* 41(1), *87.*

Table 1 Reactor Runs

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 $\label{eq:2.1} \frac{1}{2} \int_{\mathbb{R}^3} \frac{1}{\sqrt{2}} \left(\frac{1}{\sqrt{2}} \int_{\mathbb{R}^3} \frac{1}{\sqrt{2}} \left(\frac{1}{\sqrt{2}} \int_{\mathbb{R}^3} \frac{1}{\sqrt{2}} \int_{\mathbb{R}^3} \frac{1}{\sqrt{2}} \left(\frac{1}{\sqrt{2}} \int_{\mathbb{R}^3} \frac{1}{\sqrt{2}} \int_{\mathbb{R}^3} \frac{1}{\sqrt{2}} \int_{\mathbb{R}^3} \frac{1}{\sqrt{2}} \int_{\mathbb{R}^3$

 $\mathcal{L}^{\text{max}}_{\text{max}}$

Table 2 **Tetralin Soluble** Fraction **of Resid**

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Table **4** Conversion **of** thermal hydroprocessing *of* resid

a sa kabila ng pangalang ng pang
Mga pangalang ng pa

 \mathcal{L}_{max} , where \mathcal{L}_{max} and \mathcal{L}_{max}

Table *5* **Conversion of catalytic hydroprocessing** of **resid**

 $\label{eq:2} \frac{1}{2\pi\sqrt{2}}\frac{1}{\sqrt{2}}\left(\frac{1}{\sqrt{2}}\right)^{2} \left(\frac{1}{2}\right)^{2} \left(\frac{1}{2}\right)^{2} \left(\frac{1}{2}\right)^{2} \left(\frac{1}{2}\right)^{2}$

 \sim

Table *6:* **VPO results**

 \sim

* *MWs* **determined using dichlorobenzene as solvent, results likely too high.**

Table 7 Attribute distributions used to construct a molecular representation for coal resid.

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