

SYNTHESIS AND PROPERTIES OF LANTHANIDE-EXCHANGED
PREYSSLER'S HETEROPOLYANIONS

Conf-941144--78

MARK R. ANTONIO, J. MALINSKY, AND L. SODERHOLM

Chemistry Division, Argonne National Laboratory, 9700 S. Cass Avenue, Argonne, IL
60439-4831, USA

ABSTRACT

Na^+ in the Preyssler heteropolytungstate anion $[\text{NaP}_5\text{W}_{30}\text{O}_{110}]^{14-}$ can be exchanged for a trivalent lanthanide ion. The potential significance of this new class of lanthanide heteropolyanions relates to their applications in catalysis science. This view follows from the fact that Keggin heteropolyanions and their free acids are used as heterogeneous solid catalysts and homogeneous solution catalysts. We describe synthetic conditions that lead to the incorporation of Ce^{3+} and Pr^{3+} within the Preyssler anion, and the coprecipitation of Ce^{3+} and the Preyssler anion. Initial studies indicate that the latter, coprecipitated, material deserves study for bifunctional catalytic activity.

INTRODUCTION

Heteropolyanions are discrete clusters, with sizes varying from 50-200 atoms, that can accept substantial reduction in their overall net charge without decomposition.[1-3] As such, these polynuclear anions serve to bridge the gap between the chemistry of small molecular clusters with discrete, atomic-like energy levels, on the one hand, and covalent solids with delocalized, bandlike energy levels, on the other. The prototypical heteropoly oxoanions have the 1:12 Keggin structure, $[\text{PM}_{12}\text{O}_{40}]^{3-}$ ($\text{M} \equiv \text{Mo}^{6+}, \text{W}^{6+}$). Upon electrochemical reduction with one or more electrons, the anion is transformed into an intensely colored, heteropoly blue $[\text{PM}_{12}\text{O}_{40}]^{n-}$, $n=4, 5, 6$. In view of their remarkable redox behavior, heteropolyanions have received considerable attention, both experimentally and theoretically, from both academic and industrial research groups. Principal interest stems from the potential technological spin-offs to catalysis science. Heteropolymolybdates and -tungstates, for example, show activity as acid as well as oxidation and ammoxidation catalysts.[3-8] Heteropolyanions and their free acids, i.e., heteropolyacids, are used as both heterogeneous solid catalysts and homogeneous solution catalysts.

Between 1968 and 1971, four new types of heteropolyanion clusters with rare earth (RE) elements were reported:[9-12] (1) decatungstometalates, $[\text{REW}_{10}\text{O}_{36}]^{n-}$; (2) dodecamolybdometalates, $[\text{REM}_{12}\text{O}_{42}]^{n-}$; (3) bis(undecatungsto)metalates $[\text{RE}(\text{PW}_{11}\text{O}_{39})_2]^{n-}$; (4) bis(heptadecatungsto)metalates, $[\text{RE}(\text{P}_2\text{W}_{17}\text{O}_{61})_2]^{n-}$. Examples of the complexes prepared to date are summarized in Table I. More recently, the preparation of a number of bis(undecatungstocuprate)lanthanates, $[\text{RE}(\text{CuW}_{11}\text{O}_{39})_2]^{17-}$, as well as the germanium and boron derivatives, $[\text{RE}(\text{GeW}_{11}\text{O}_{39})_2]^{13-}$ and $[\text{RE}(\text{BW}_{11}\text{O}_{39})_2]^{15-}$, have been described,[13-14] Table I. Such heteropolyoxometalate clusters containing rare earth ions with dual valence states, such as Ce, Pr, Tb (3+ and 4+) or Sm, Eu, Tm, Yb (2+ and 3+), and multivalent (3+, 4+, 5+, 6+) transition metal ions of Group VIB are of particular interest. These can have unusual redox properties—facile, reversible reduction-oxidation—that may be exploited for catalysis. In fact, a recent study described the behavior of $[\text{Ce}^{4+}\text{W}_{10}\text{O}_{36}]^{8-}$, $[\text{Nd}^{3+}\text{W}_{10}\text{O}_{36}]^{9-}$, and $[\text{Sm}^{3+}\text{W}_{10}\text{O}_{36}]^{9-}$ for the homogenous, selective oxidation of cyclohexanol with H_2O_2 and for H_2O_2 decomposition.[15]

DISCLAIMER

This report was prepared as an account of work sponsored by an agency of the United States Government. Neither the United States Government nor any agency thereof, nor any of their employees, makes any warranty, express or implied, or assumes any legal liability or responsibility for the accuracy, completeness, or usefulness of any information, apparatus, product, or process disclosed, or represents that its use would not infringe privately owned rights. Reference herein to any specific commercial product, process, or service by trade name, trademark, manufacturer, or otherwise does not necessarily constitute or imply its endorsement, recommendation, or favoring by the United States Government or any agency thereof. The views and opinions of authors expressed herein do not necessarily state or reflect those of the United States Government or any agency thereof.

DISCLAIMER

Portions of this document may be illegible in electronic image products. Images are produced from the best available original document.

Table I. Rare-earth (*RE*) element-containing heteropolyanions with the cluster charge (*n*) and the *RE* ions reported to form these structures.

Heteropolyanion	Charge (<i>n</i>)	<i>RE</i>
[<i>REW</i> ₁₀ O ₃₆] ^{<i>n</i>-}	9	3+ (Y, La-Yb)
	8	4+ (Ce, Th, U)
[<i>REMo</i> ₁₂ O ₄₂] ^{<i>n</i>-}	9	3+ (Ce)
	8	4+ (Ce, Th, U, Np)
	7	5+ (U)
[<i>RE(PW</i> ₁₁ O ₃₉) ₂] ^{<i>n</i>-}	11	3+ (La, Ce, Pr, Nd, Gd)
	10	4+ (Ce, U)
	9	5+ (U)
[<i>RE(SiW</i> ₁₁ O ₃₉) ₂] ^{<i>n</i>-}	13	3+ (Ce, Sm, Eu, Tb, Ho)
	12	4+ (Ce, U)
	11	5+ (U)
[<i>RE(CuW</i> ₁₁ O ₃₉) ₂] ^{<i>n</i>-}	17	3+ (La-Gd, Dy, Ho, Yb)
[<i>RE(P</i> ₂ <i>W</i> ₁₇ O ₆₁) ₂] ^{<i>n</i>-}	17	3+ (Ce, Tb, Dy, Ho)
	16	4+ (Ce, Tb, Th, U)
[<i>RESb</i> ₉ <i>W</i> ₂₁ O ₈₆] ^{<i>n</i>-}	16	3+ (La-Gd, Dy, Yb)
[<i>REP</i> ₅ <i>W</i> ₃₀ O ₁₁₀] ^{<i>n</i>-}	12	3+ (Y, Nd-Lu)
	11	4+ (U)

Among the large number and variety of heteropolyoxolanthanate clusters, it is most common that lacunary (i.e., unsaturated) polyanions act as ligands for the *RE* ions. In contrast, there are just two heteropolyanions of composition [*NaSb*₉*W*₂₁O₈₆]¹⁸⁻ and [*NaP*₅*W*₃₀O₁₁₀]¹⁴⁻ that encapsulate rare earth ions through Na⁺ exchange,[16,17] the latter of which is known as the Preyssler anion, PA. A single-crystal X-ray structure study reveals that it contains Na⁺ within a central cylindrical cavity formed by the cyclic arrangement of five -*PW*₆O₂₂- groups assembled with D_{5h} symmetry.[18] Under relatively rigorous conditions, Na⁺ can be exchanged with *RE*³⁺ (*RE*≡Y, Nd-Lu) to form heteropolyoxotungstate anions [*REP*₅*W*₃₀O₁₁₀]¹²⁻.[17] The significance and interest of these *RE*-exchanged heteropolyanions relates to their possible utilization in separation and catalysis sciences. For example, the cavity through the Preyssler anion bears some resemblance to the tunnels connecting cages in microporous aluminophosphate molecular sieves and zeolites. In view of this, *RE*-ion exchanged Preyssler anions may have properties in common with *RE*-ion exchanged zeolites, which are in widespread use as hydrocarbon cracking catalysts and may have some impact on the catalytic reduction of NO.[19] The catalytic efficiency of the Preyssler anion for the selective oxidation of H₂S by molecular oxygen in aqueous solution was the subject of a recent investigation by Harrup and Hill.[20]

This report describes new preparative chemistry for the exchange of Na⁺ in [*NaP*₅*W*₃₀O₁₁₀]¹⁴⁻ with Ce³⁺ and Pr³⁺. We have also performed initial studies on the coprecipitation of Ce³⁺ and the Preyssler anion. The motivation for study of the latter PA-supported Ce³⁺ complex stems from the preparation of rhodium and iridium polyoxometalates of [*SiW*₁₂O₄₀]⁴⁻ and [*PMo*₁₂O₄₀]³⁻. These have been described as bifunctional catalysts with diverse reaction chemistry, including olefin isomerization, hydrogenation, hydroformylation, and C-H activation processes.[21] Similarly, polyoxoanion-supported iridium and ruthenium (pre)catalysts have been prepared and characterized.[22,23] The cyclic voltammetry (CV) and *RE* L₃-edge XANES (X-ray

absorption near edge structure) of $[REP_5W_{30}O_{110}]^{12-}$ and the Ce^{3+} salt of $[NaP_5W_{30}O_{110}]^{14-}$ are described herein.

EXPERIMENTS

The white, crystalline Preyssler salt, $K_{12.5}Na_{1.5}[NaP_5W_{30}O_{110}] \cdot 15H_2O$, was prepared according to the method of ref 17. The Ce^{3+} and Pr^{3+} -exchange reactions were based upon the following equation:



In a typical experiment for $n=1$, a colorless warm ($60^\circ C$) solution of $Ce(NO_3)_3 \cdot 6H_2O$ (53 mg, 0.121 mmole, dissolved in 3 cm^3 of H_2O) was added dropwise with stirring to a warm ($60^\circ C$) solution of the Preyssler salt (1 g, 0.121 mmol, dissolved in 12 cm^3 H_2O). The resulting clear, colorless solution was sealed in a Parr 4746 Teflon-lined digestion vessel and heated at $165^\circ C$ in a Lindberg/Blue M crucible furnace for 48 h. Upon cooling, 4 g of solid KCl was added to the clear yellow solution to precipitate a fine pale yellow powder, which was collected on No. 42 Whatman filter paper, rinsed with ca. 10 cm^3 of ice-cold water and dried in air. The same procedure was followed for $n=2$ in the equation above. Yields of 0.86–0.98 g (80–90% based upon $[NaP_5W_{30}O_{110}]^{14-}$ and an estimated empirical formula $K_{12}[CeP_5W_{30}O_{110}] \cdot 54H_2O$) were obtained. The Ce^{3+} exchange reaction occurs with both one and two equivalents of Ce^{3+} as either $Ce(NO_3)_3 \cdot 6H_2O$ or $CeCl_3 \cdot 7H_2O$. The exchange of Na^+ with Pr^{3+} also occurs in aqueous solutions of $Pr(NO_3)_3 \cdot 6H_2O$ using the conditions described above. After heating at $165^\circ C$ for 48 h, the clear colorless solution is worked up to produce a white powder.

The following method was used to coprecipitate Ce^{3+} with $[NaP_5W_{30}O_{110}]^{14-}$. A clear, colorless aqueous solution containing $Ce(NO_3)_3 \cdot 6H_2O$ ($n=2$) and the Preyssler salt (prepared exactly as above) was evaporated to dryness at $135^\circ C$ for 2 h in a drying oven. This coprecipitation produced a pastel yellow solid.

Ce and Pr L_3 -edge XANES data were collected at ambient temperature on beam line 4-3 (using a Si<220> double crystal monochromator) at SSRL and line X-23A2 (using a Si<311> double crystal monochromator) at NSLS. The vertical entrance slit width was set to 1 mm. The fluorescence signal, I_f , was detected by use of an ion chamber fluorescent detector (The EXAFS Co.) filled with argon and without a fluorescence filter. The powder samples were mounted at the conventional 45° incident— 45° exit configuration in a helium purged sample box. Nitrogen was used to monitor the incident X-ray intensity, I_0 . Due to the low concentrations of Ce and Pr, ca. 1.6 wt %, in the exchanged Preyssler anions, the fluorescence XANES, I_f/I_0 , is not vitiated by self-absorption effects. The XANES data for the model compounds were obtained by the electron-yield, I_e/I_0 , method. The normalization of the X-ray absorption data to a unit edge jump was performed according to conventional methods.[24]

Cyclic voltammetry data were obtained with a BAS 100B/W electrochemical analyzer, a BAS C-2 voltammetry cell stand, and BAS electrodes: 3.0 mm diameter glassy carbon working electrode (MF-2012); platinum wire auxiliary electrode (MW-1032); Ag/AgCl reference electrode with vycor tip (MF-2063). The electrolyte (1 M H_2SO_4) was prepared from 99.9999% H_2SO_4 (Alfa) and 18M Ω -cm water from a Millipore MILLI-Q™ water system. CV scans of the neat electrolyte, which was sparged and blanketed with nitrogen, did not reveal any evidence of electroactive impurities in the potential windows of interest, -0.60 to 0.20 V vs. Ag/AgCl. The scan rates were -100 mV/s and the anion concentrations were ca. 1 mM.

RESULTS AND DISCUSSION

The CVs of the Ce^{3+} - and Pr^{3+} -exchanged Preyssler anions are identical. As shown in Figure 1 for $[\text{CeP}_5\text{W}_{30}\text{O}_{110}]^{12-}$, the CV data exhibit 5 reversible redox waves attributable to $\text{W}(6+/5+)$ and the formation of heteropoly blues. For comparison, the CV for the parent Preyssler anion is also shown in Figure 1. The differences in the CVs between the RE^{3+} -exchanged anions and the parent anion serve to confirm the exchange of Na^+ with Ce^{3+} and Pr^{3+} . In fact, the CVs for the Ce^{3+} - and Pr^{3+} -exchanged anions shown here are identical to the CVs for all the other RE^{3+} -exchanged anions, except Eu^{3+} . [17,25] Figure 1 also shows the CV for the yellow material obtained by coprecipitation of Ce^{3+} and $[\text{NaP}_5\text{W}_{30}\text{O}_{110}]^{14-}$. It has features that are similar to those in the CV of the Preyssler ion itself. Although additional studies are clearly required, the similarities of the CVs suggest that the Ce^{3+} cations are not interacting in a chemically significant way with the $[\text{NaP}_5\text{W}_{30}\text{O}_{110}]^{14-}$ anion framework. This situation of non-interacting cations and anions is similar to that exemplified by the bifunctional catalysts $[(\text{Ph}_3\text{P})_2\text{Rh}(\text{CO})]_4[\text{SiW}_{12}\text{O}_{40}]$ and $[(\text{Ph}_3\text{P})_2\text{IrH}_2]_3[\text{PMo}_{12}\text{O}_{40}]$. [21]

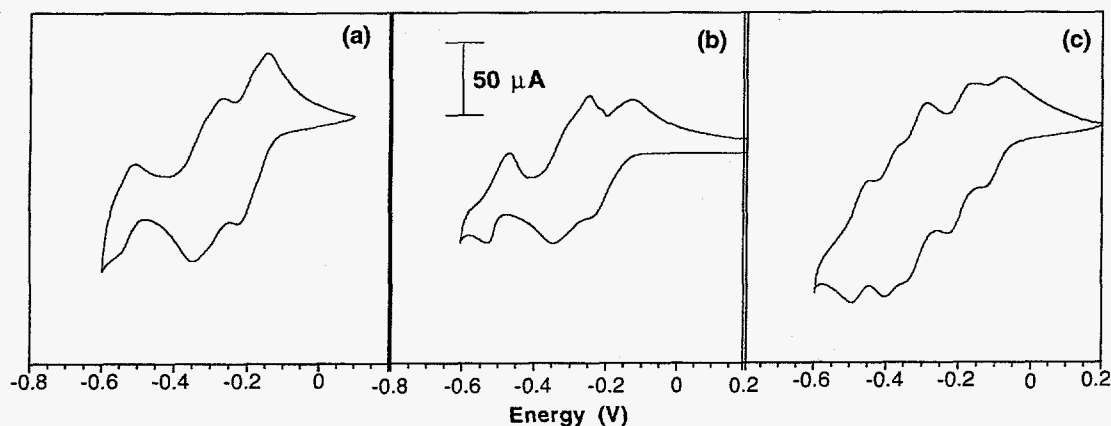


Figure 1. Cyclic voltammograms of (a) $\text{K}_{12.5}\text{Na}_{1.5}[\text{NaP}_5\text{W}_{30}\text{O}_{110}] \cdot 15\text{H}_2\text{O}$ (b) $\text{K}_{12.5}\text{Na}_{1.5}[\text{NaP}_5\text{W}_{30}\text{O}_{110}] \cdot 15\text{H}_2\text{O}$ coprecipitated with 2 equivalents of $\text{Ce}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$ (c) $\text{K}_{12}[\text{CeP}_5\text{W}_{30}\text{O}_{110}] \cdot n\text{H}_2\text{O}$ in aqueous 1 M H_2SO_4 electrolytes.

The Ce and Pr L_3 -edge XANES for the solid salts from the Ce and Pr-exchange reactions and the Ce coprecipitation are shown in Figure 2. The intense single edge resonances at 5721.6 eV (Ce L_3 XANES) and 5965.4 eV (Pr L_3 XANES) are due to $2p \rightarrow 5d$ electronic transitions. The edge profiles and positions found here are typical of those observed in L_3 -edge XANES for trivalent Ce and Pr compounds, [26] such as for $\text{Ce}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$ and $\text{Pr}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$, see Figure 2. These spectra stand in sharp contrast to the L -edge XANES for compounds containing quadrivalent cerium and

praseodymium, e.g., CeO_2 and PrO_2 . For such materials, two well-resolved edge resonances are observed.[27] The XANES for the Ce^{3+} -exchanged-PA solid salt shown here is similar to that for the cerium-exchanged PA prepared according to the original method from aqueous Ce^{4+} [17]. In fact, the Ce XANES for the $\text{K}_{12}[\text{CeP}_5\text{W}_{30}\text{O}_{110}] \cdot n\text{H}_2\text{O}$ solid salt of Figure 2 is identical to that from the aqueous solution of $[\text{CeP}_5\text{W}_{30}\text{O}_{110}]^{12-}$ prepared from Ce^{4+} . [28] Whereas cerium is trivalent in the exchange products obtained with either Ce^{3+} or Ce^{4+} and in view of the fact that the corresponding CV data are identical, the differences between the XANES of the two solid state salts are surprising. They may suggest the presence of a Ce^{3+} impurity in the original preparations or, alternatively, different locations for Ce^{3+} within the cylindrical cavity of the Preyssler anion.[28] Additional studies to resolve this issue are in progress.

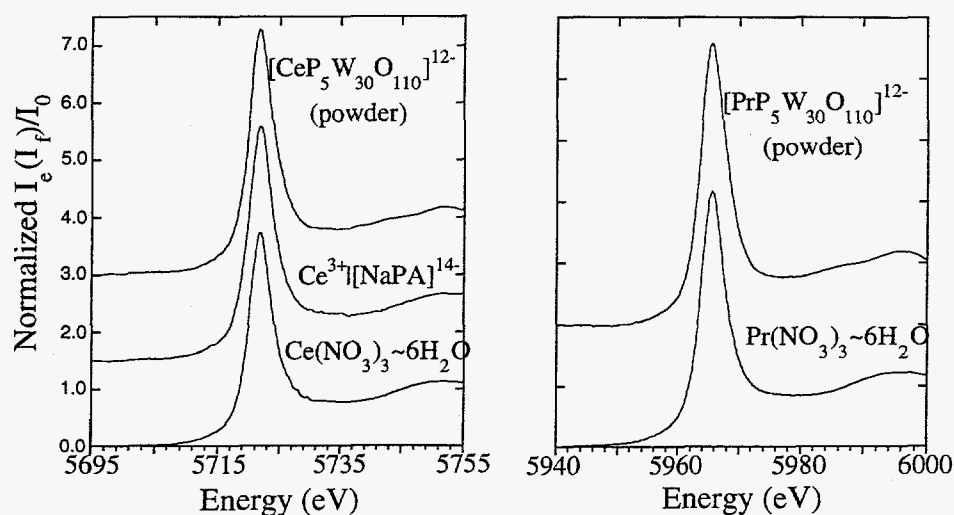


Figure 2. Solid-state L_3 -edge XANES for (a) $\text{K}_{12}[\text{CeP}_5\text{W}_{30}\text{O}_{110}] \cdot n\text{H}_2\text{O}$, $\text{K}_{12.5}\text{Na}_{1.5}[\text{NaP}_5\text{W}_{30}\text{O}_{110}] \cdot 15\text{H}_2\text{O}$ coprecipitated with 2 equivalents of $\text{Ce}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$; $\text{Ce}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$; (b) $\text{K}_{12}[\text{PrP}_5\text{W}_{30}\text{O}_{110}] \cdot n\text{H}_2\text{O}$, $\text{Pr}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$.

CONCLUSIONS

The exchange of Na^+ in the Preyssler anion, $[\text{NaP}_5\text{W}_{30}\text{O}_{110}]^{14-}$, with Ce^{3+} and Pr^{3+} was investigated. Under extreme conditions of temperature (165°C) and time (48 h), the direct exchange of Na^+ for Ce^{3+} and Pr^{3+} occurs from aqueous solutions of $\text{Ce}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$ (or $\text{CeCl}_3 \cdot 7\text{H}_2\text{O}$) and $\text{Pr}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$. The Preyssler anion was also precipitated with trivalent cerium. The cyclic voltammetry of the Ce and Pr exchange products confirm the Na^+ ion exchange, and the L_3 -edge XANES show that Ce and Pr are trivalent. Although no catalysis screening studies were performed here, this work provides a new entry to the synthesis of heteropolyanions with potential catalytic activity.

ACKNOWLEDGMENTS

This work was supported by the U.S. DOE, BES—Chemical Sciences, under contract No. W-31-109-ENG-38.

REFERENCES

1. M.T. Pope, *Heteropoly and Isopoly Oxometalates* (Springer-Verlag, Berlin, 1983).
2. M.T. Pope, in *Comprehensive Coordination Chemistry*, Vol. 3, edited by G. Wilkinson, R.D. Gillard and J.A. McCleverty (Pergamon Press, New York, 1987), pp. 1023-1058.
3. M.T. Pope and A. Müller, *Angew. Chem. Int. Ed. Engl.* **30**, 34 (1991).
4. M. Misono, in *Polyoxometalates: From Platonic Solids to Anti-Retroviral Activity*, edited by M.T. Pope and A. Müller (Kluwer Academic, Dordrecht, The Netherlands, 1994) pp. 255-265; R.G. Finke, *ibid.*, pp. 267-280; J.H. Grate, D.R. Hamm and S. Mahajan, *ibid.*, pp. 281-305; R. Neumann, *ibid.*, pp. 307-313; E. Cadot, C. Marchal, M. Fournier, A. Teze and G. Herve, *ibid.*, pp. 315-326; E. Papaconstantinou, A. Ioannidis, A. Hiskia, P. Argitis, D. Dimotikali and S. Korres, *ibid.*, pp. 327-335; C.L. Hill, G.-S. Kim, C.M. Prosser-Mccartha and D. Judd, *ibid.*, pp. 359-371.
5. N. Mizuno and M. Misono, *J. Molecular Catalysis* **86**, 319 (1994).
6. M. Misono, *Catal. Rev.—Sci. Eng.* **29**, 269 (1987).
7. I.V. Kozhevnikov and K.I. Matveev, *Russian Chemical Reviews* **51**, 1075 (1982)
8. D. Sattari and C.L. Hill, *J. Am. Chem. Soc.* **115**, 469 (1993)
9. R.D. Peacock and T.J.R. Weakley, *J. Chem. Soc. (A)* **1971**, 1836.
10. P. Baidala, V.S. Smurova, E.A. Torchenkova and V. I. Spitsyn, *Doklady Chemistry* **197**, 430 (1971), *Engl. Trans.*
11. E.A. Torchenkova, P. Baidala, V.S. Smurova and V.I. Spitsyn, *Doklady Chemistry* **199**, 568 (1971), *Engl. Trans.*
12. D.D. Dexter and J.V. Silverton, *J. Am. Chem. Soc.* **90**, 3589 (1968).
13. W. Qingyin, W. Enbo and L. Jingfu, *Polyhedron* **12**, 2563 (1993).
14. N. Haraguchi, Y. Okaue, Y. Isobe and Y. Matsuda, *Inorg. Chem.* **33**, 1015 (1994).
15. R. Shiozaki, H. Goto and Y. Kera, *Bull. Chem. Soc. Jpn.* **66**, 2790 (1993).
16. J. Liu, S. Liu, L. Qu, M.T. Pope and C. Rong, *Trans. Met. Chem.* **17**, 311 (1992).
17. I. Creaser, M.C. Heckel, R.J. Neitz and M.T. Pope, *Inorg. Chem.* **32**, 1573 (1993).
18. M.H. Alizadeh, S.P. Harmalker, Y. Jeannin, J. Martin-Frere and M.T. Pope, *J. Am. Chem. Soc.* **107**, 2662 (1985).
19. C. Yokoyama and M. Misono, *Bull. Chem. Soc. Jpn.* **67**, 557 (1994).
20. M.K. Harrup and C.L. Hill, Presented at the 208th National Meeting of the American Chemical Society, Washington, D.C., August 1994; paper INOR 567.
21. A.R. Siedle, R.A. Newmark, W.B. Gleason, R.P. Skarjune, K.O. Hodgson, A.L. Roe and V.W. Day, *Solid State Ionics* **26**, 109 (1988). A.R. Siedle, R.A. Newmark, K.A. Brown-Wensley, R.P. Skarjune, L.C. Haddad, K.O. Hodgson and A.L. Roe, *Organometallics* **7**, 2078 (1988). A.R. Siedle, C.G. Markell, P.A. Lyon, K.O. Hodgson and A.L. Roe, *Inorg. Chem.* **26**, 219 (1987)
22. W.G. Klemperer and B. Zhong, *Inorg. Chem.* **32**, 5821 (1993).
23. Y. Lin, K. Nomiya and R.G. Finke, *Inorg. Chem.* **32**, 6040 (1993)
24. M.R. Antonio, in *Encyclopedia of Materials Characterization: Surfaces, Interfaces, Thin Films*, edited by C. R. Brundle, C. A. Evans, Jr. and S. Wilson (Butterworth-Heinemann, Boston, 1992), p. 214.
25. M.R. Antonio, L. Soderholm, J. Muntean and G. Liu, in preparation (1994).
26. J.E. Sunstrom, IV, S.M. Kauzlarich and M.R. Antonio, *Chem. Mater.* **5**, 182 (1993).
27. Z. Hu, S. Bertram and G. Kaindl, *Phys. Rev. B*, **49**, 39 (1994).
28. M.R. Antonio and L. Soderholm, *Inorg. Chem.*, in the press (1994).