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## Metal Organic Framework Modifications of Structural Fibers

Marwan Al-Haik

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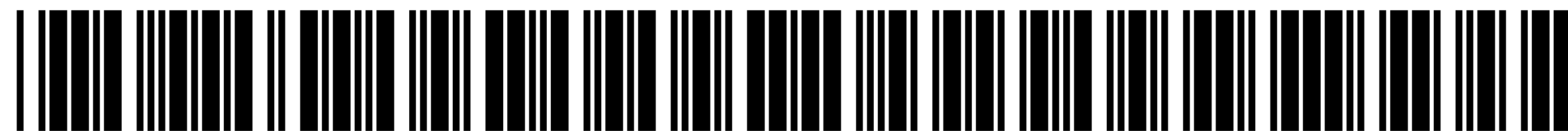
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(54) **METAL ORGANIC FRAMEWORK  
MODIFICATIONS OF STRUCTURAL FIBERS**

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*B32B 5/02* (2006.01)

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(52) **U.S. Cl.**  
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(US)

(21) Appl. No.: **17/685,003**

(57) **ABSTRACT**

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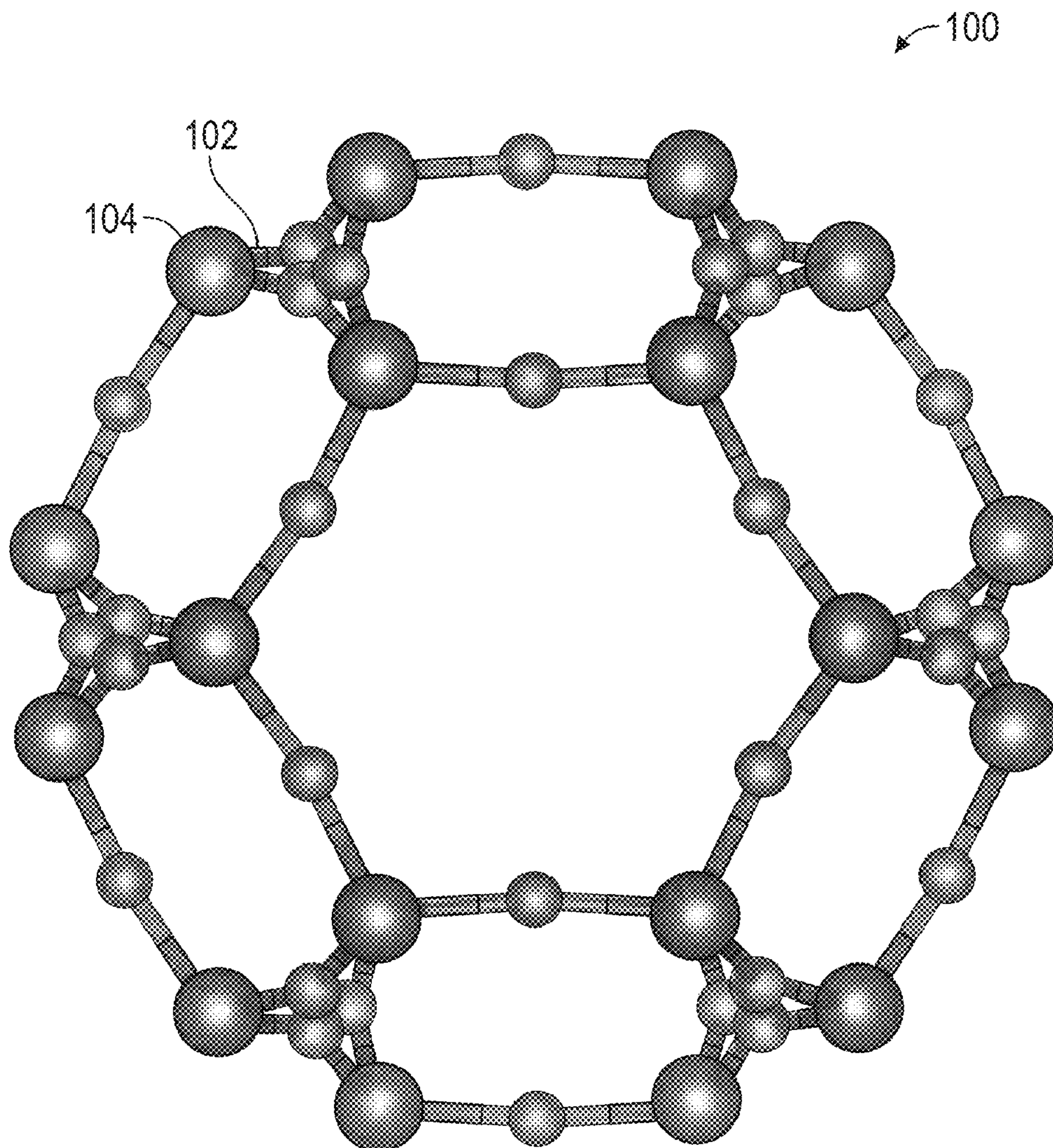
**Related U.S. Application Data**

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2, 2021.

**Publication Classification**

(51) **Int. Cl.**  
*B32B 9/04* (2006.01)  
*B32B 7/12* (2006.01)

A reinforced carbon composite can include a carbon substrate and a metal organic framework bonded to the carbon substrate. For example, a reinforced carbon composite can include a first layer, a second layer, and a resin adhered to the first layer and the second layer. The first layer can include a carbon substrate and a metal organic framework bonded to the carbon substrate. The second layer can include a carbon substrate and a metal organic framework bonded to the carbon substrate.





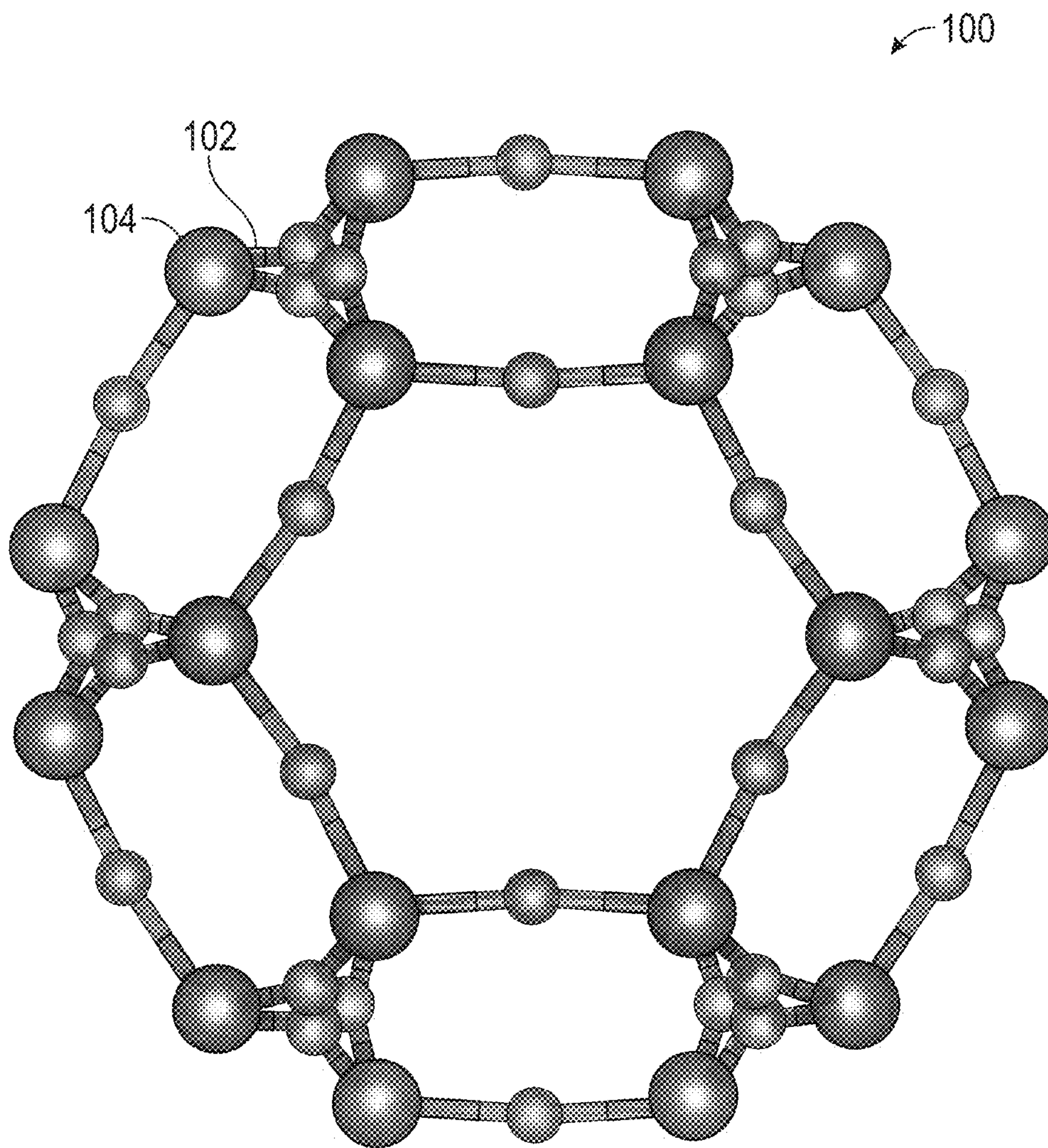


FIG. 1

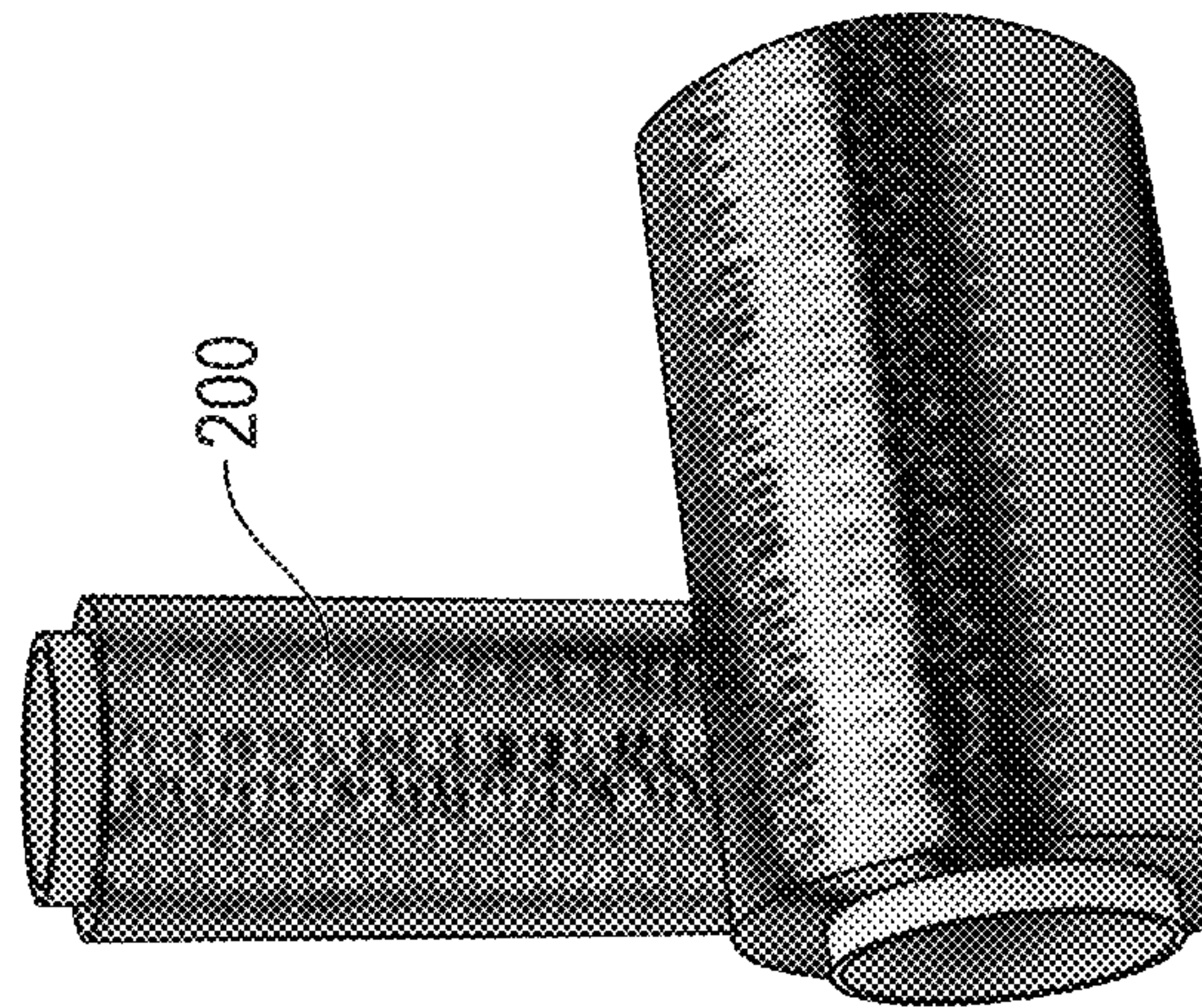
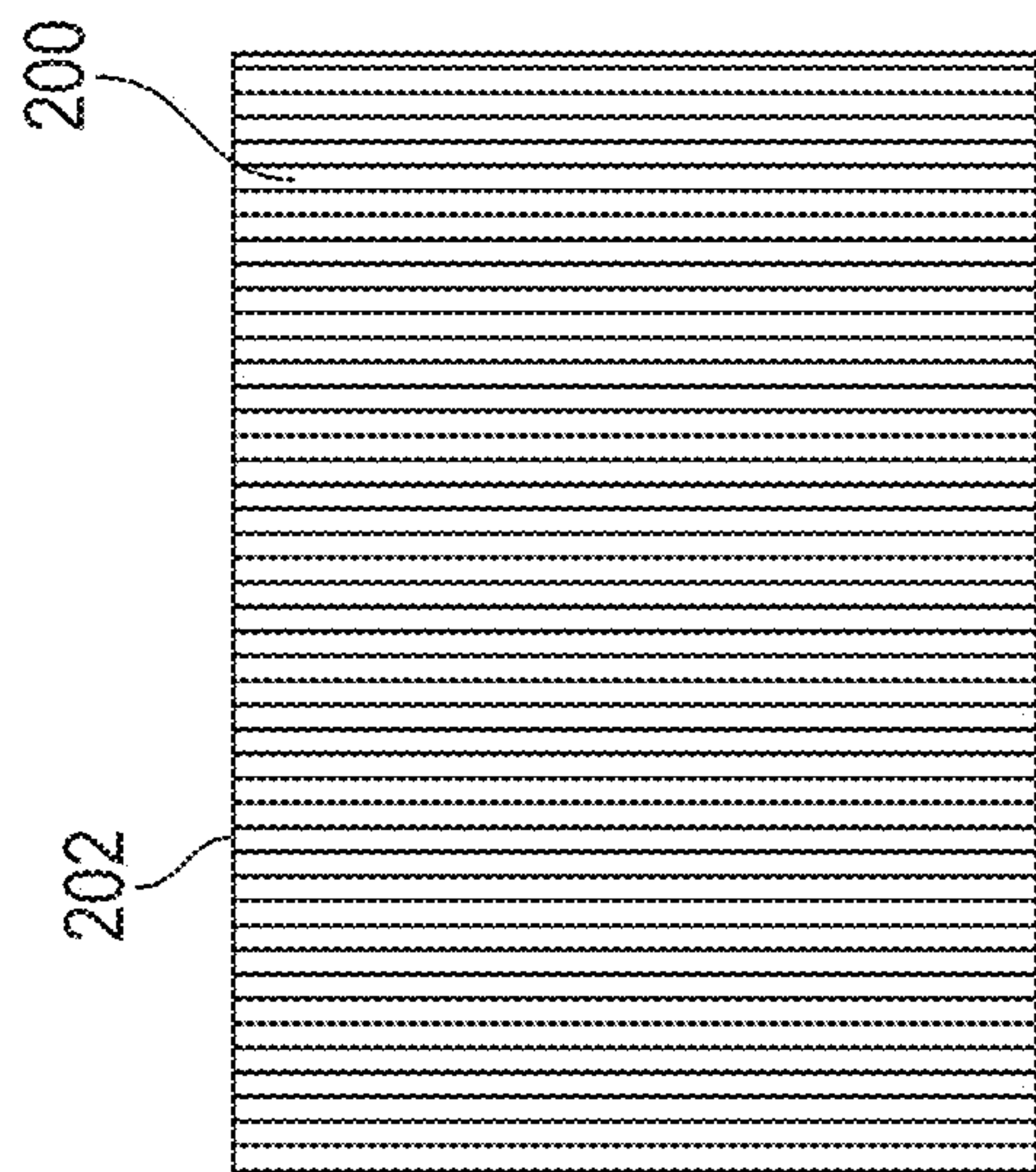
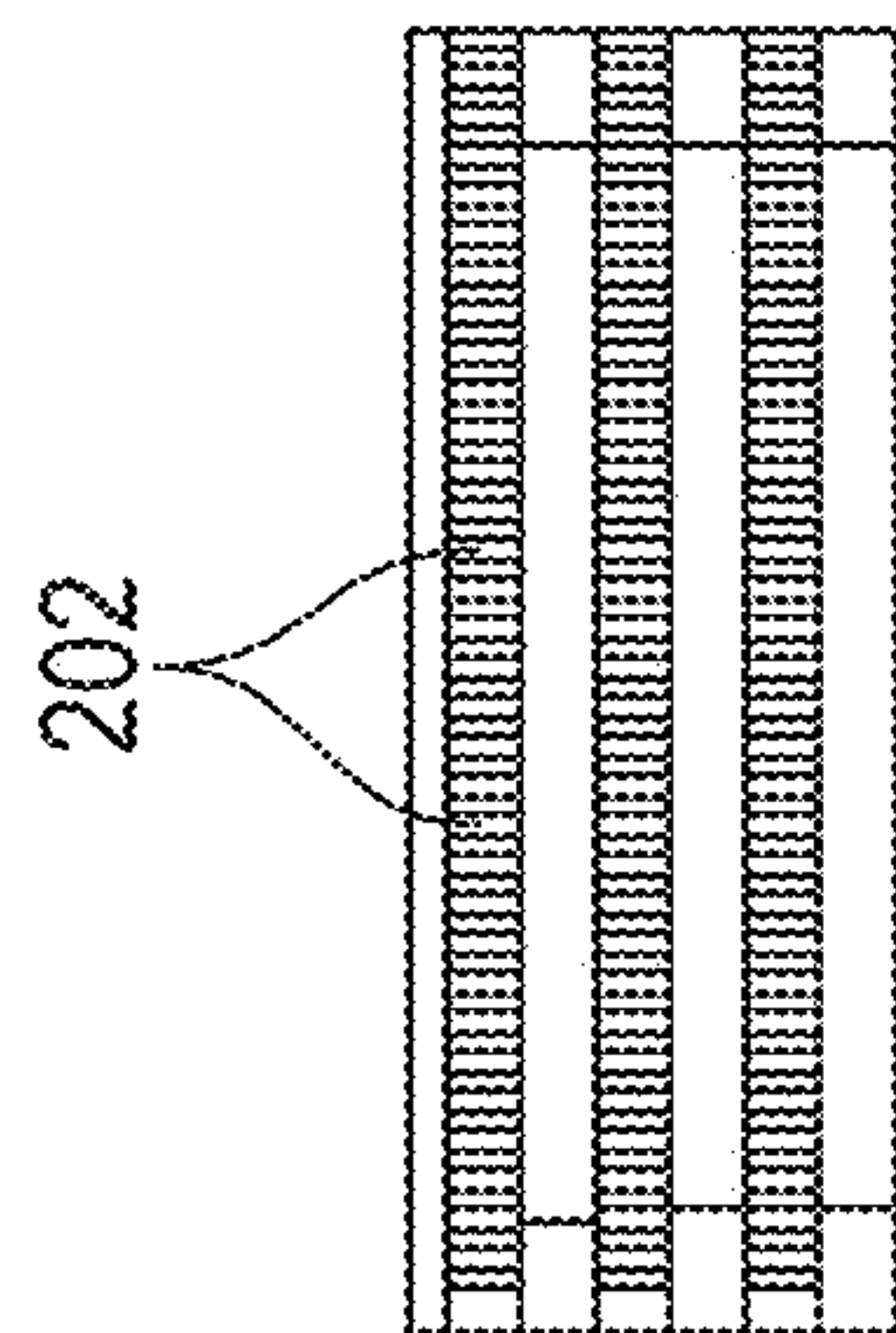


FIG. 2A



Wrap CF Around a Glass Plate  
Do Acid Treatment



Immerse 3 Plates in Solution  
with Enough Concentration Per  
Carbon Fiber Area

FIG. 2C



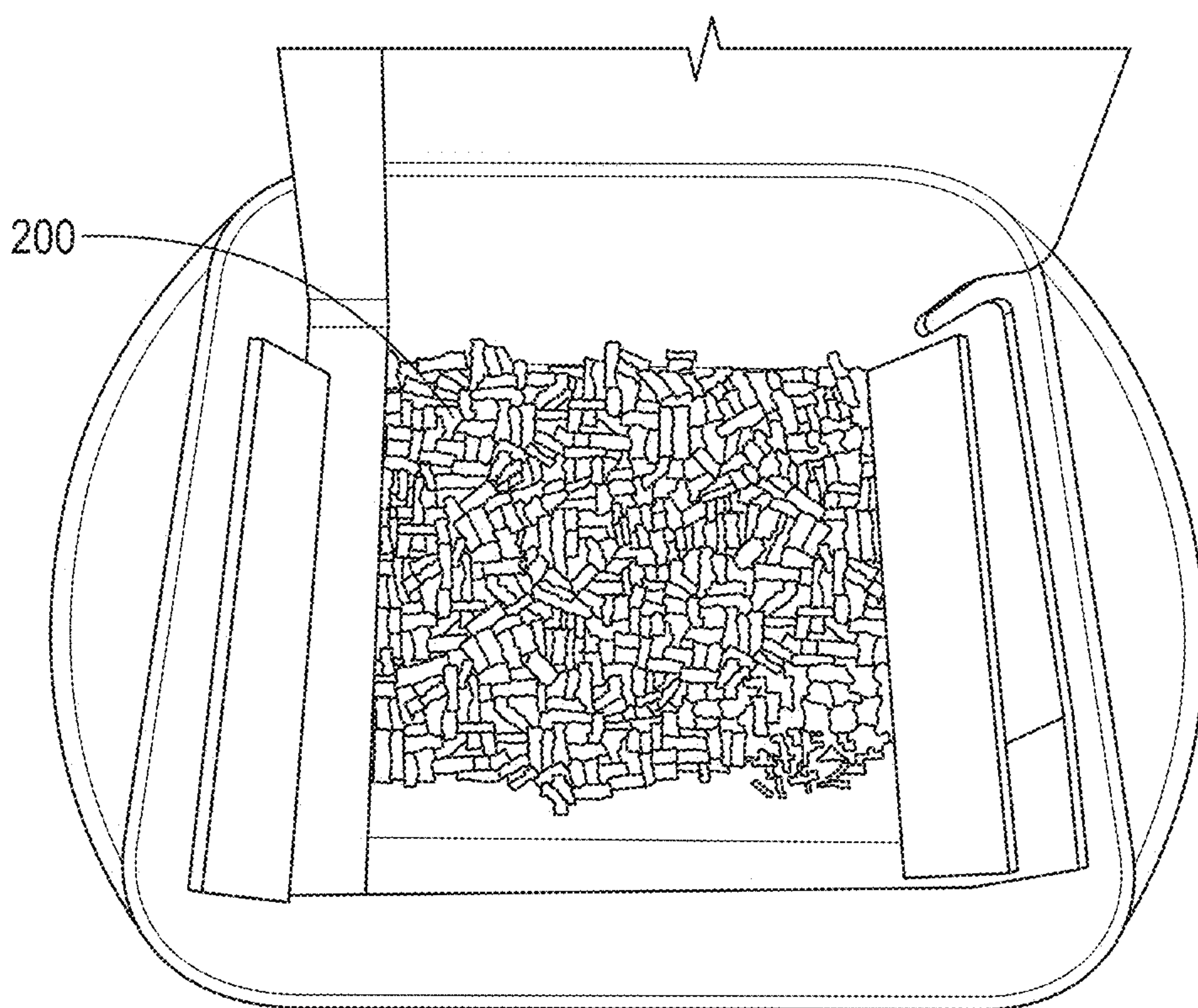


FIG. 3



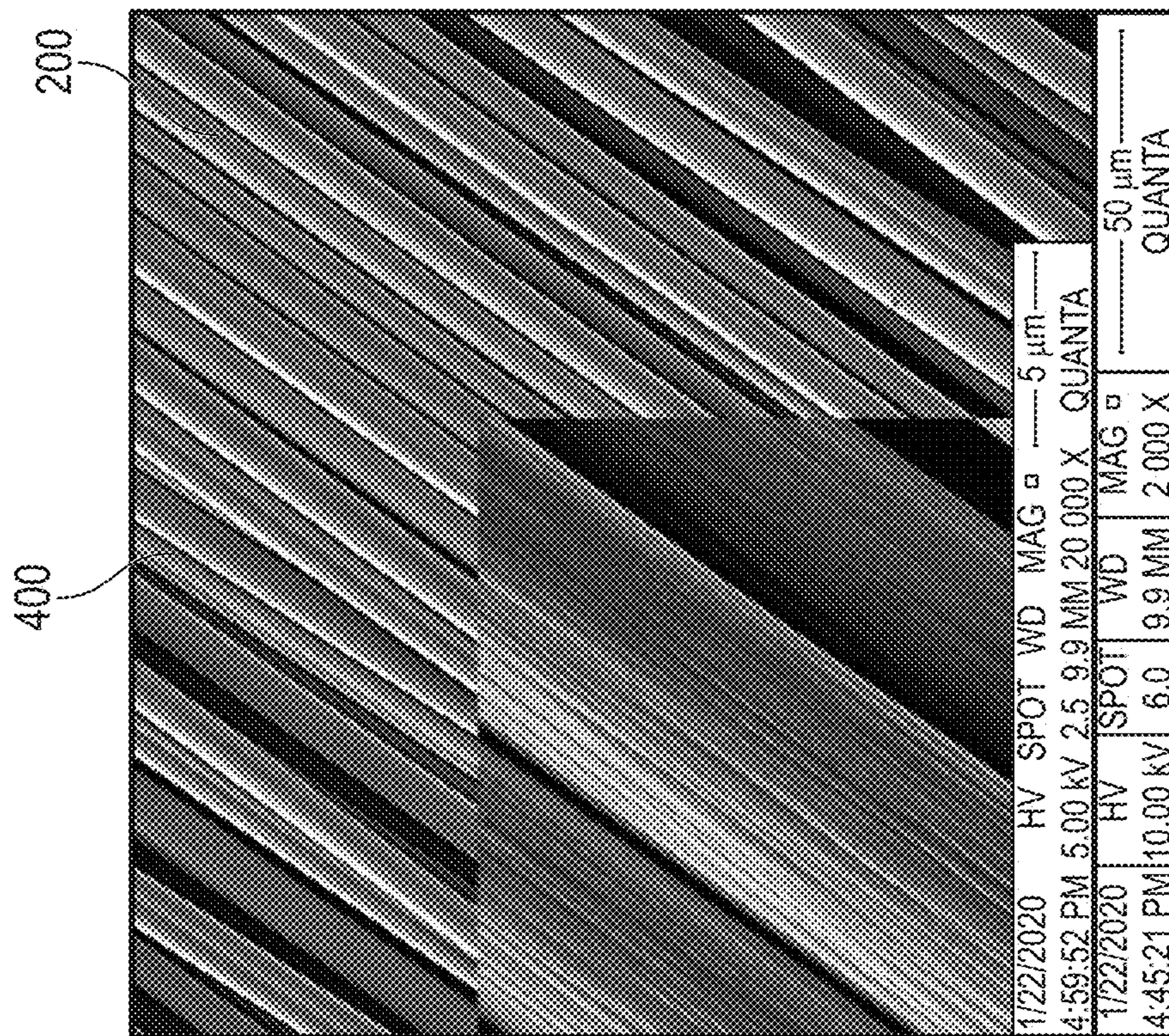


FIG. 4B

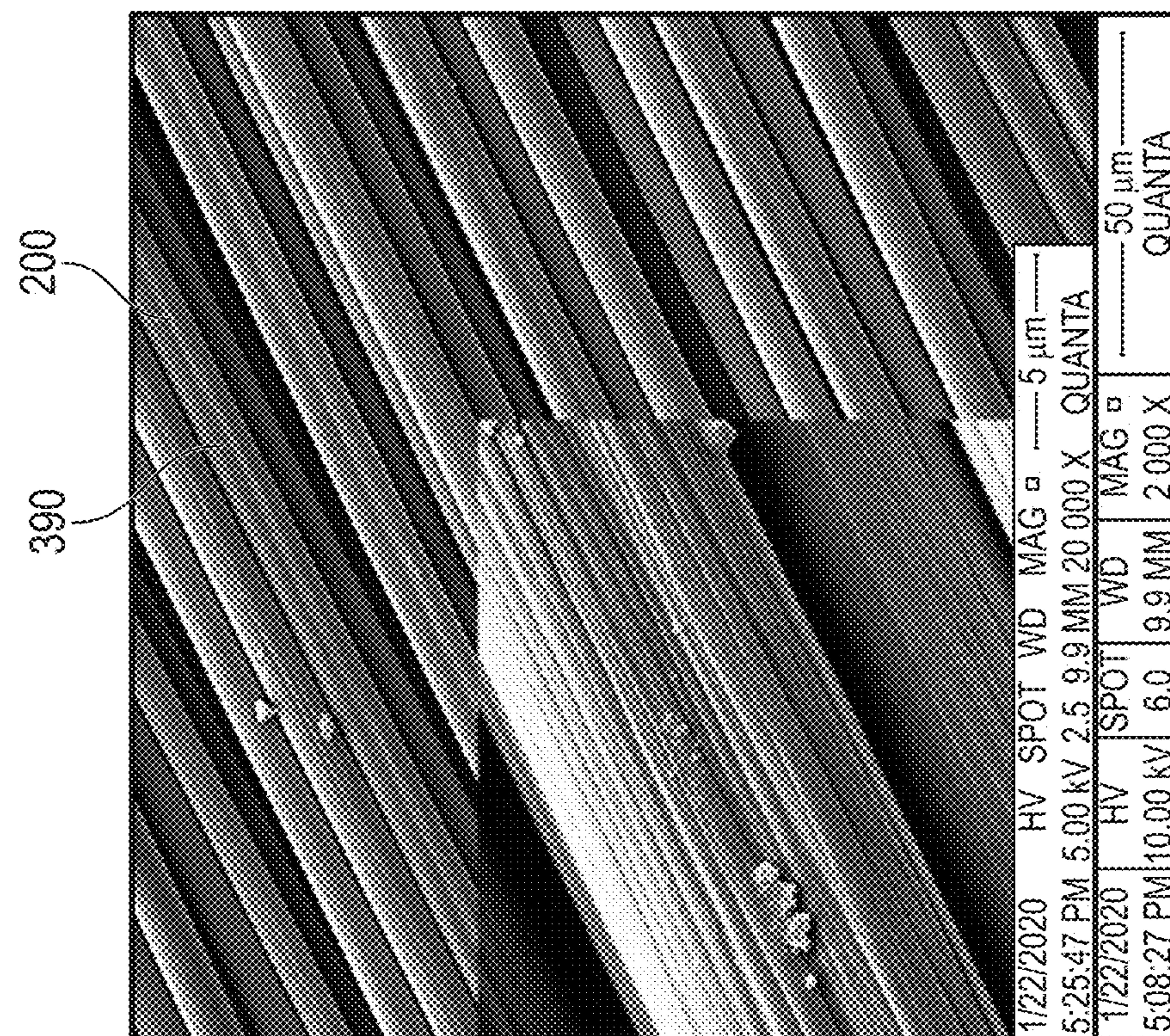


FIG. 4A



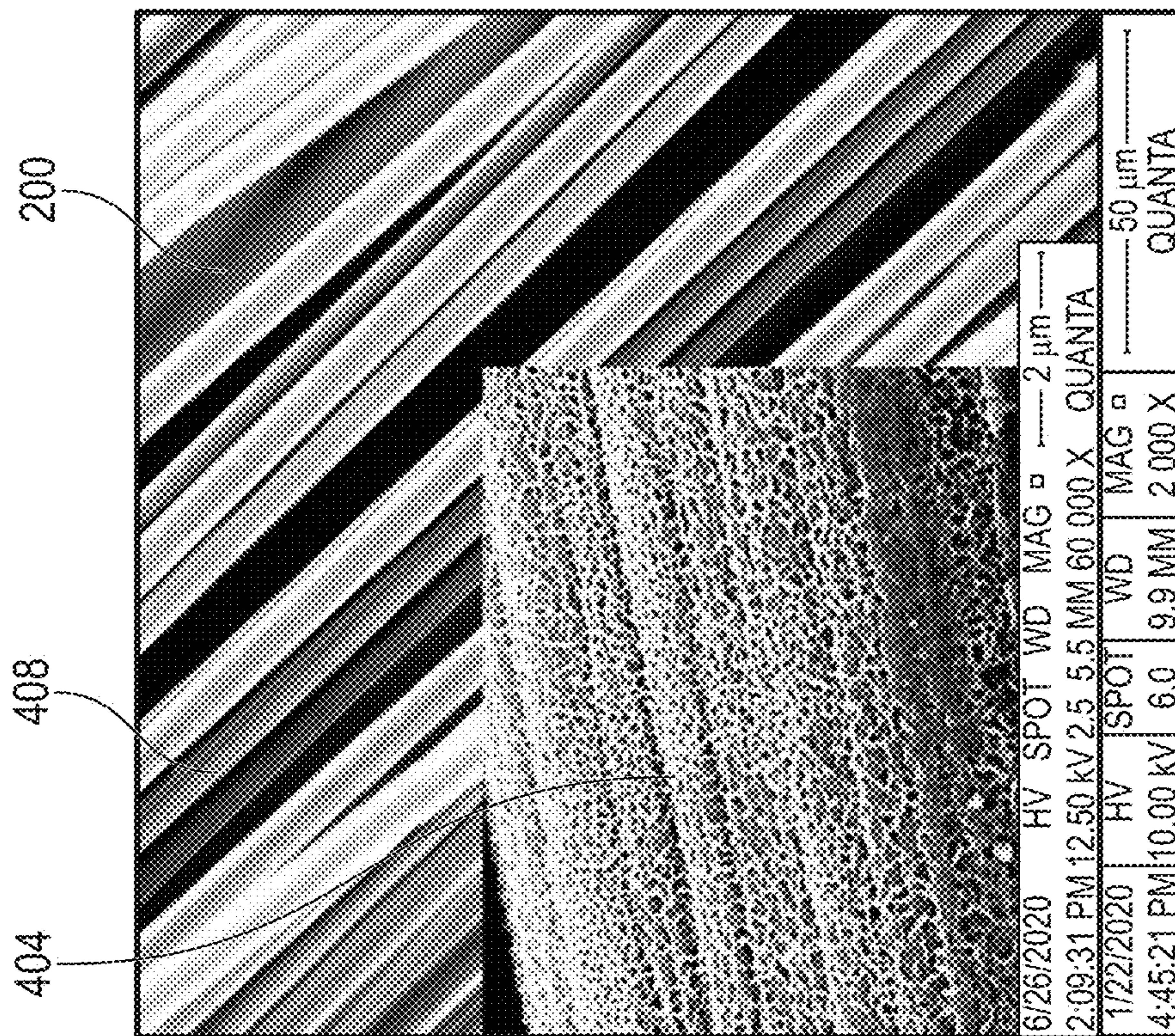


FIG. 4D

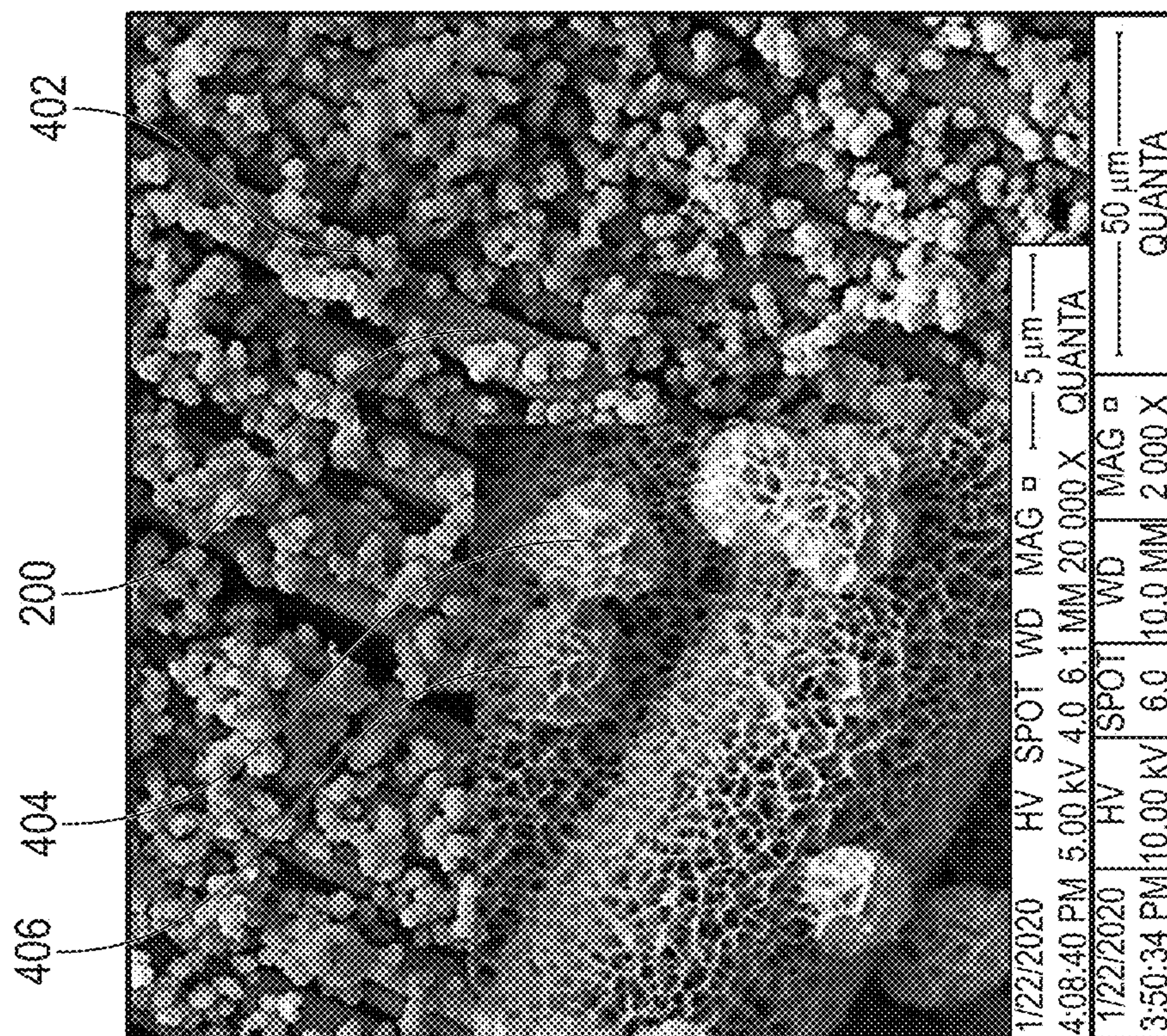


FIG. 4C



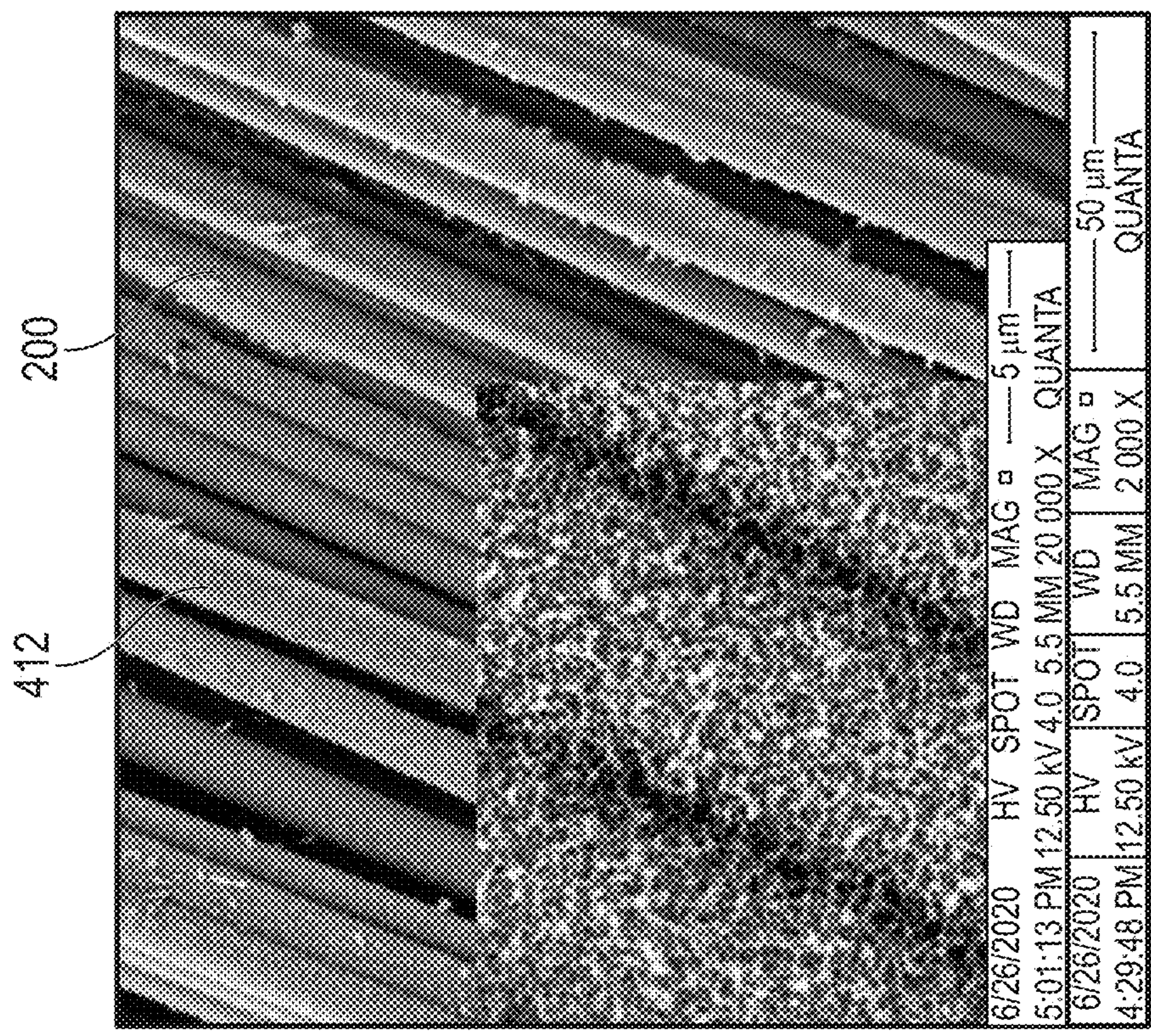


FIG. 4F

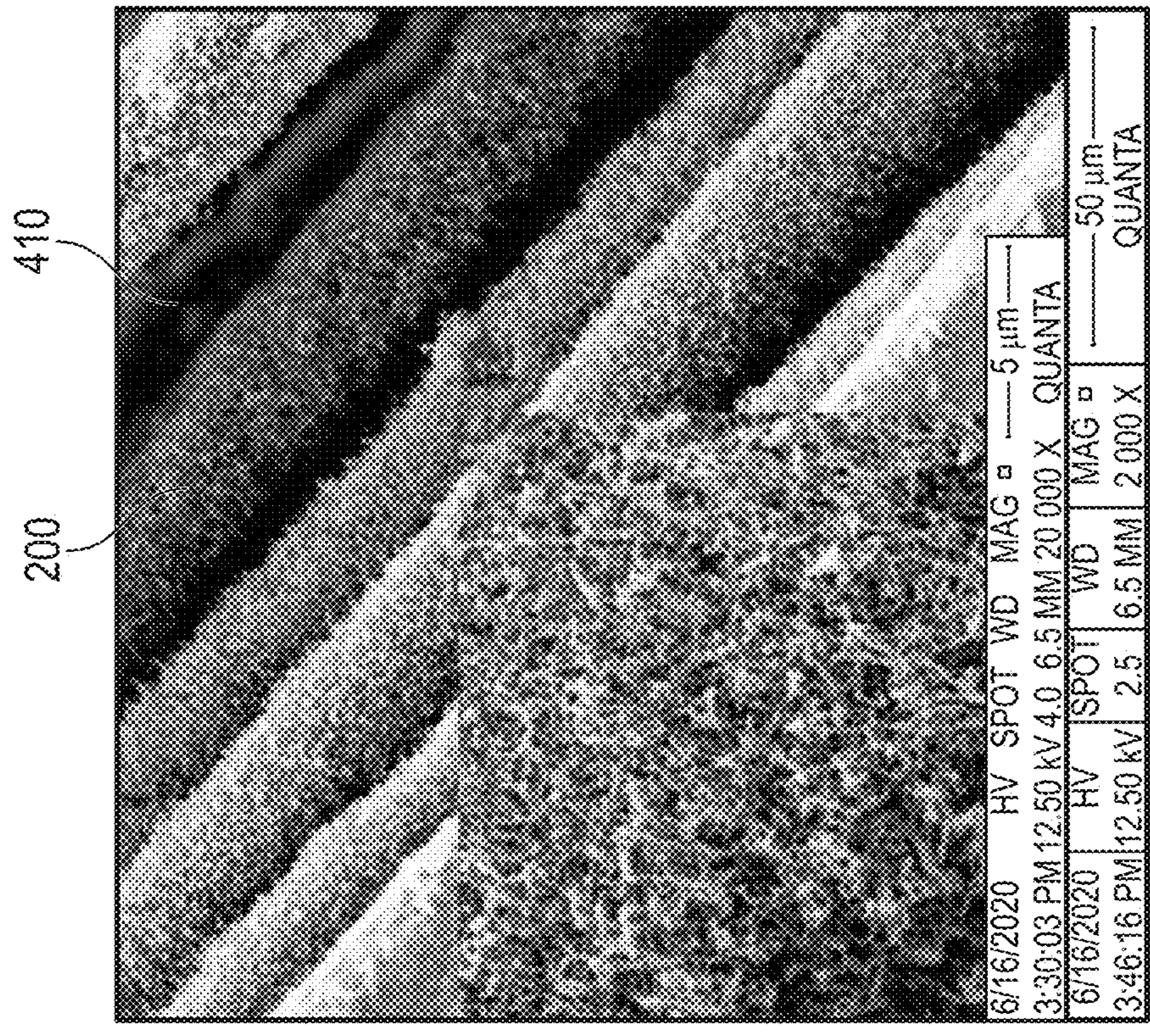


FIG. 4E



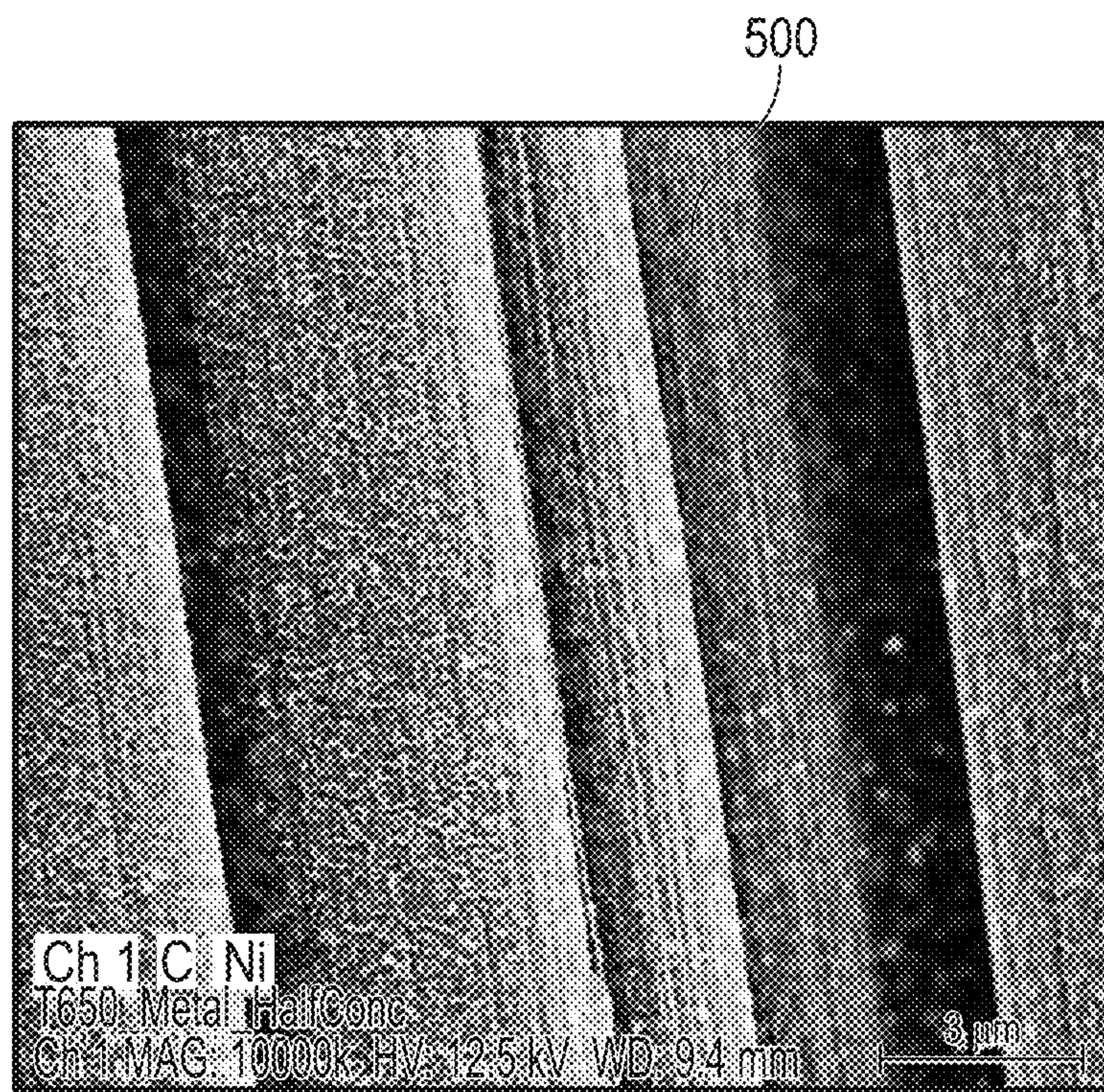


FIG. 5A



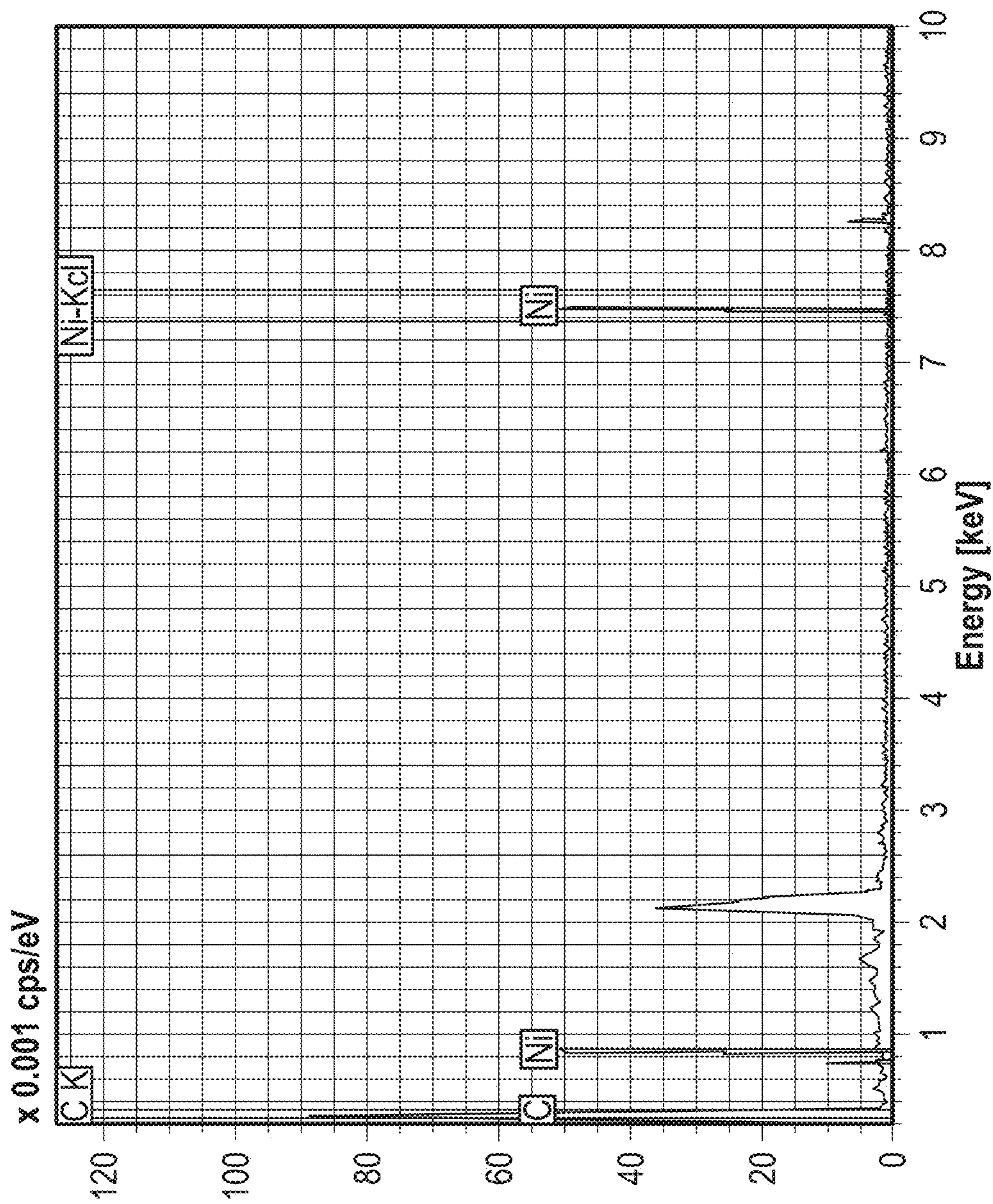


FIG. 5B



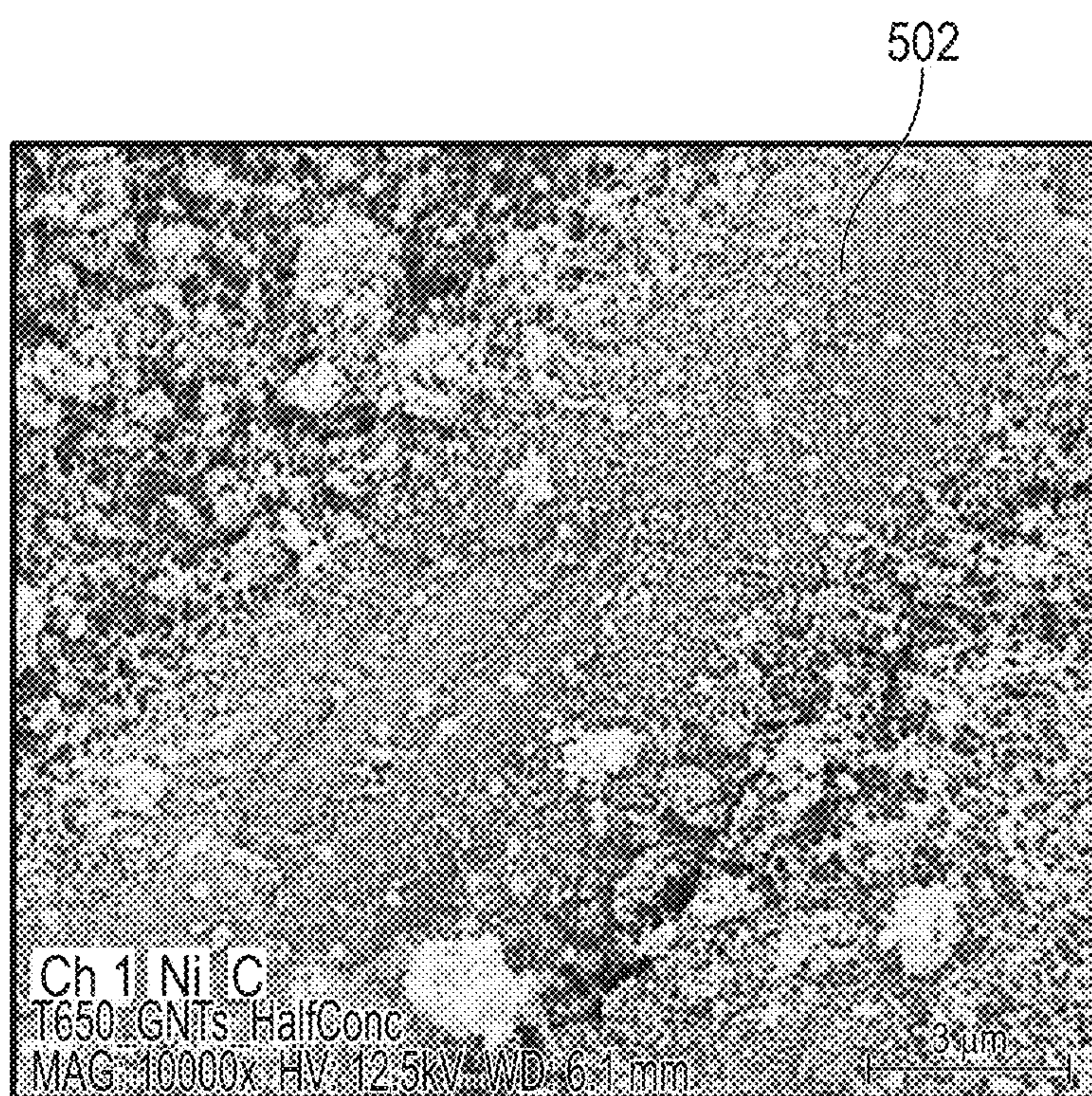


FIG. 5C



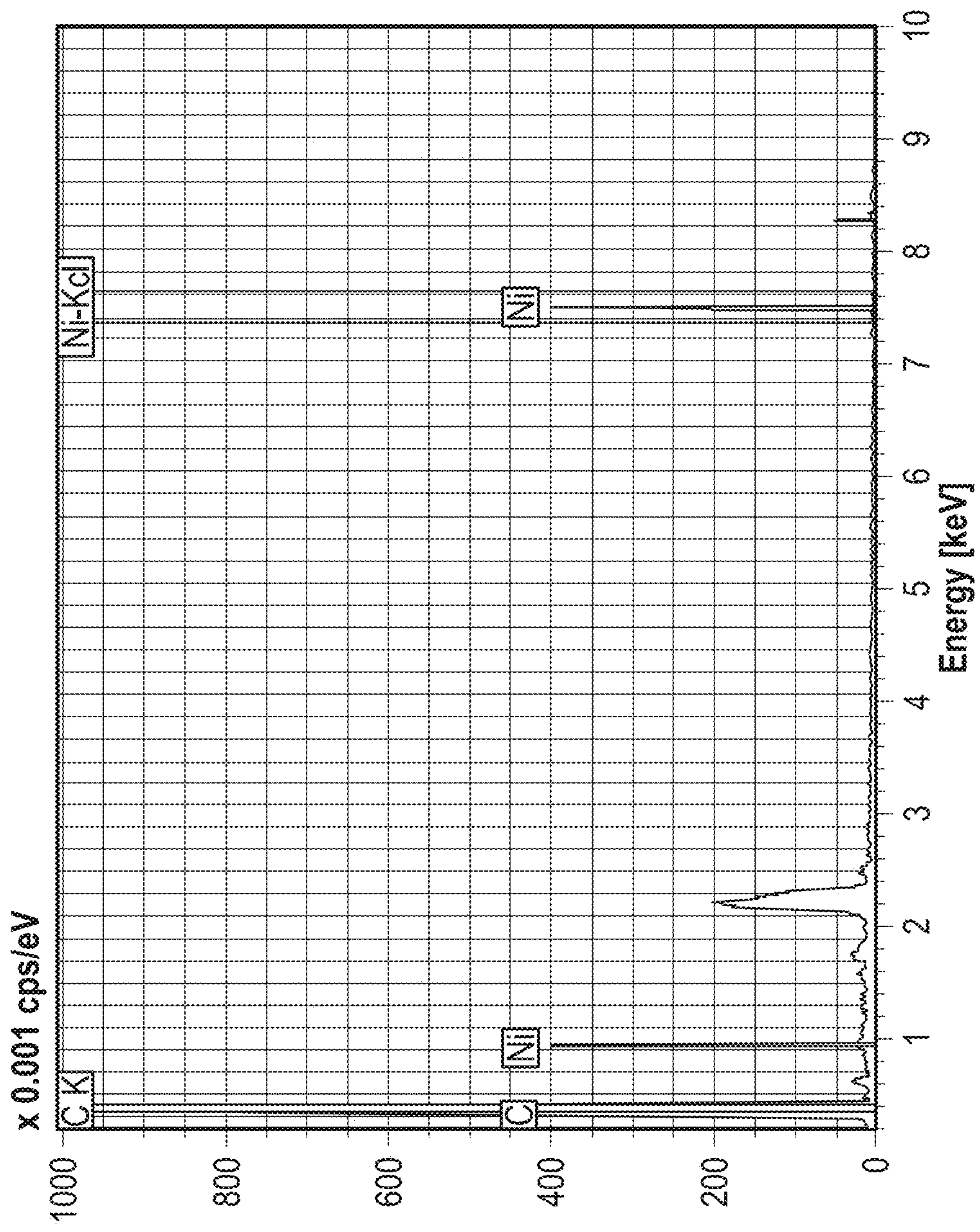


FIG. 5D



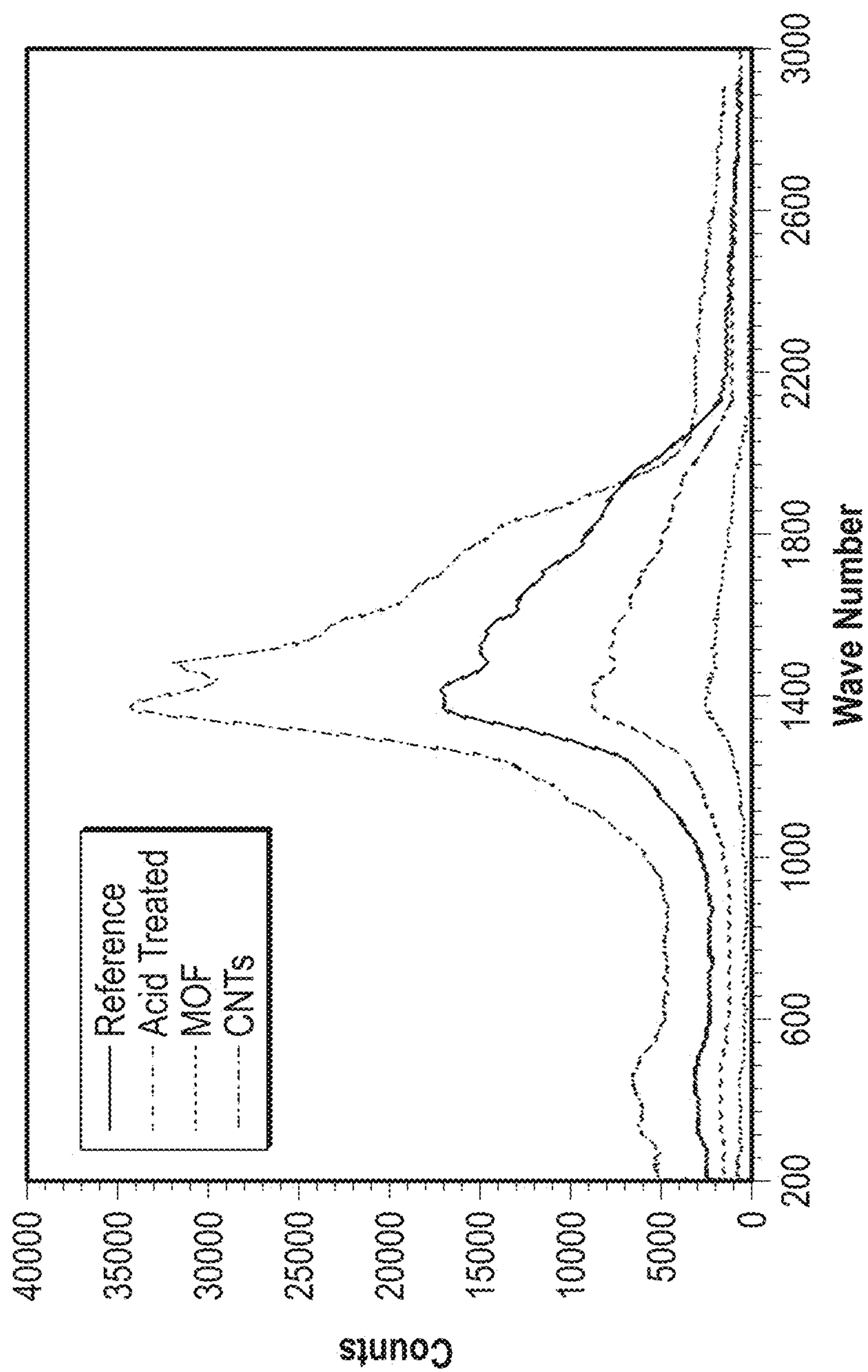


FIG. 6



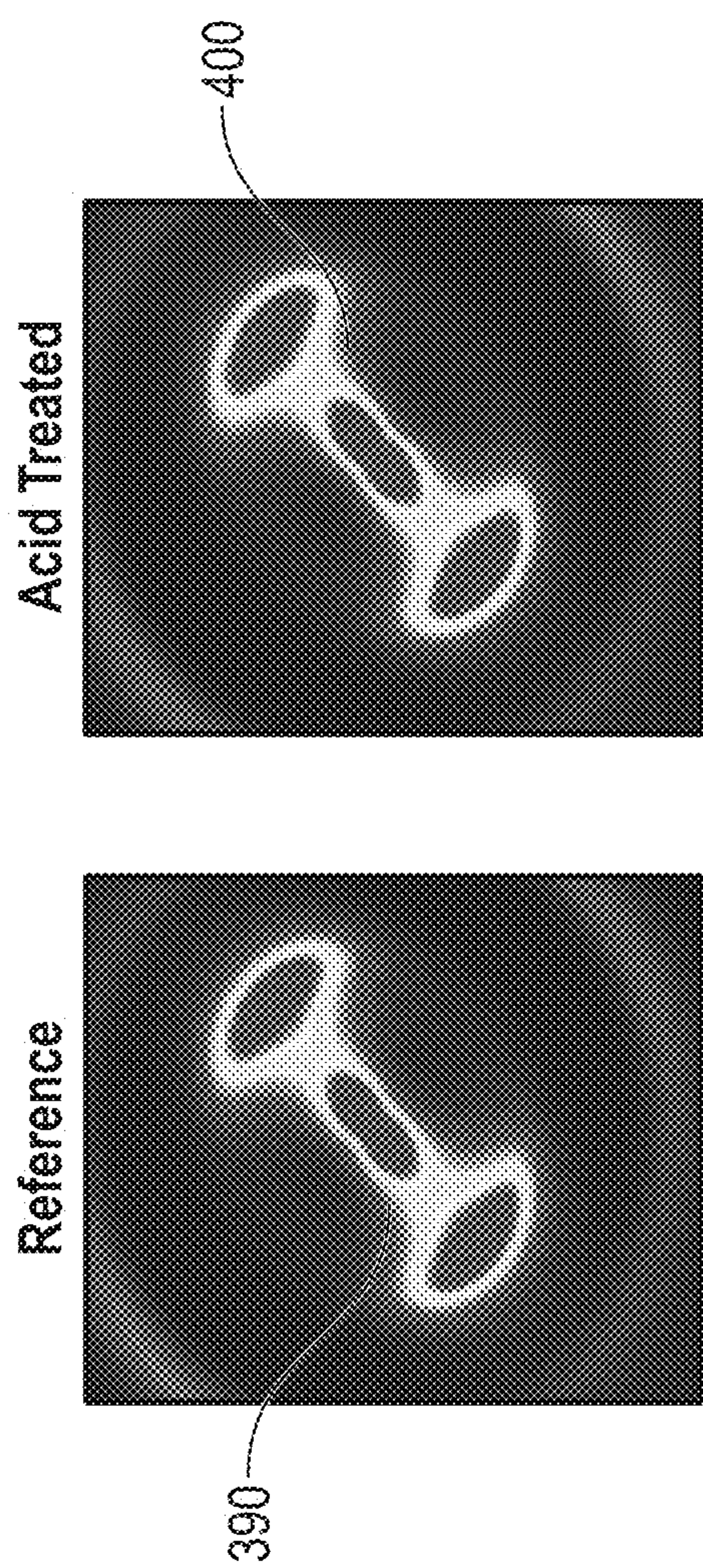


FIG. 7A

FIG. 7B

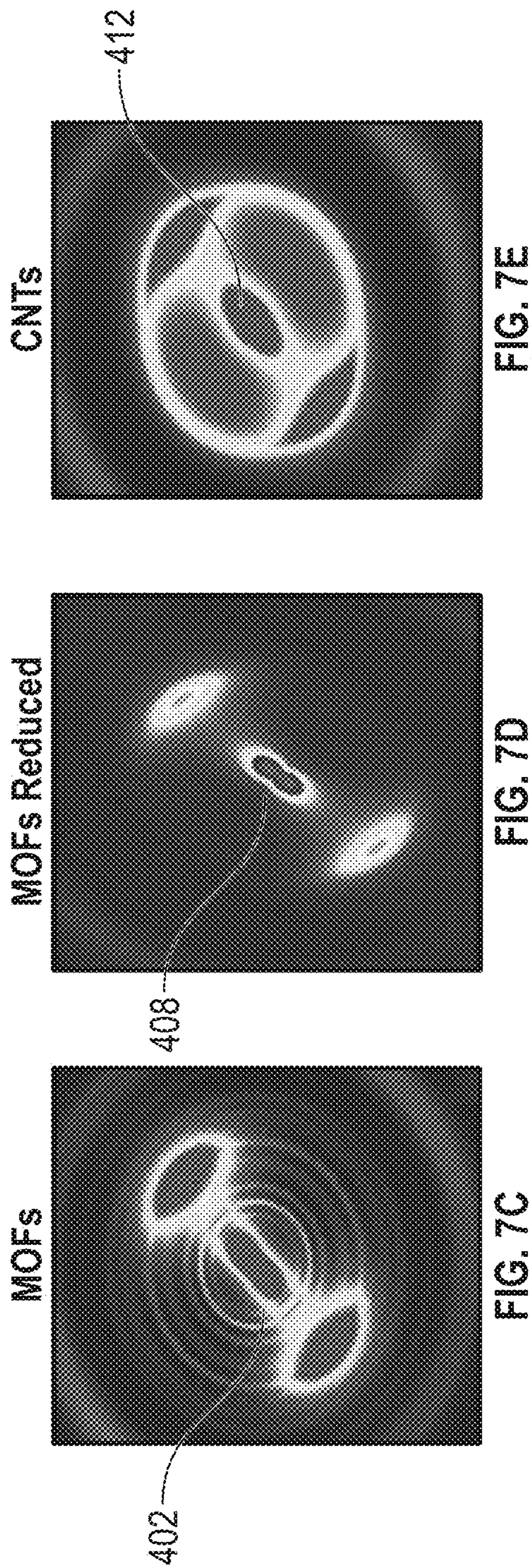


FIG. 7C

FIG. 7D

FIG. 7E



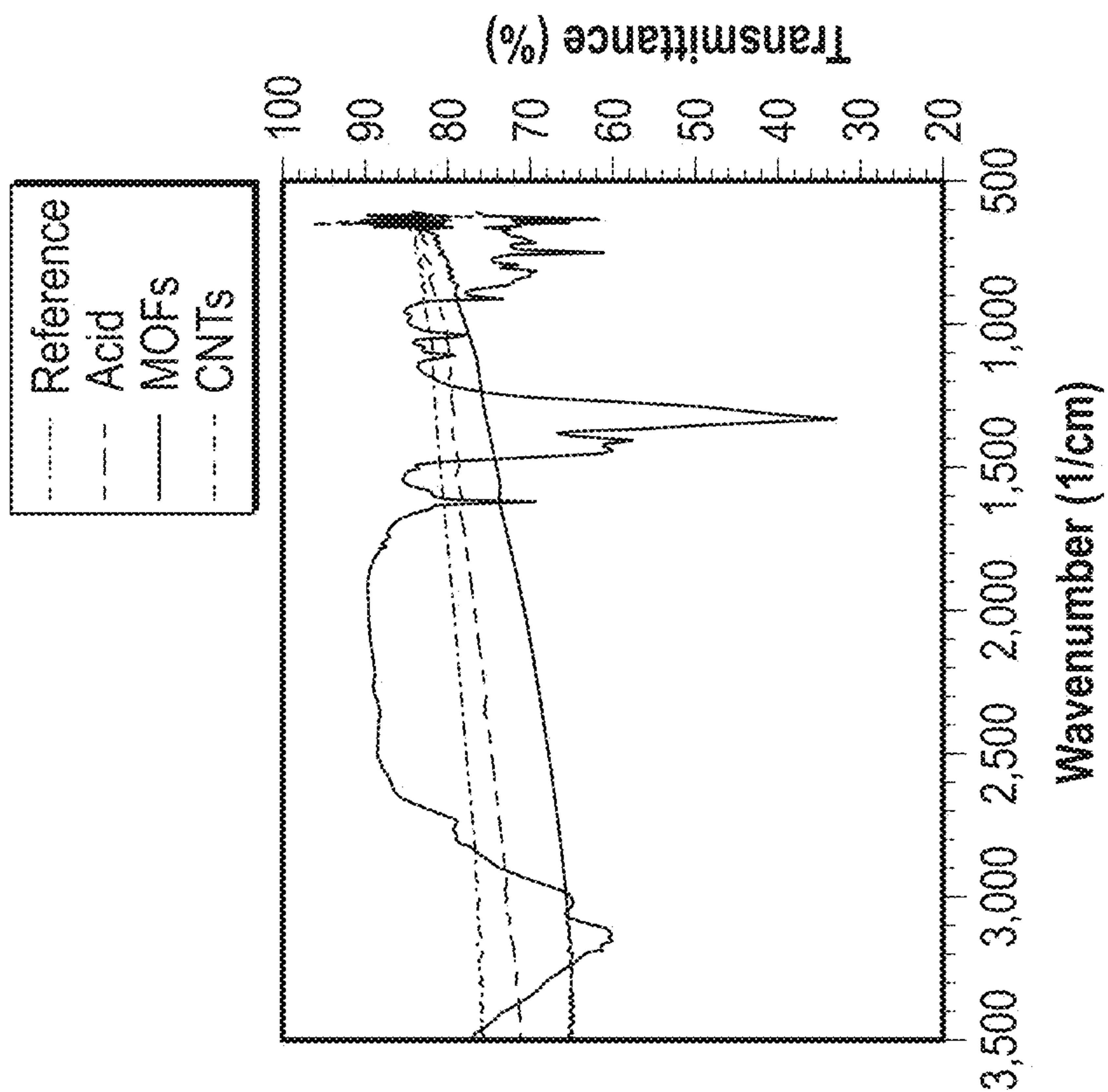


FIG. 9

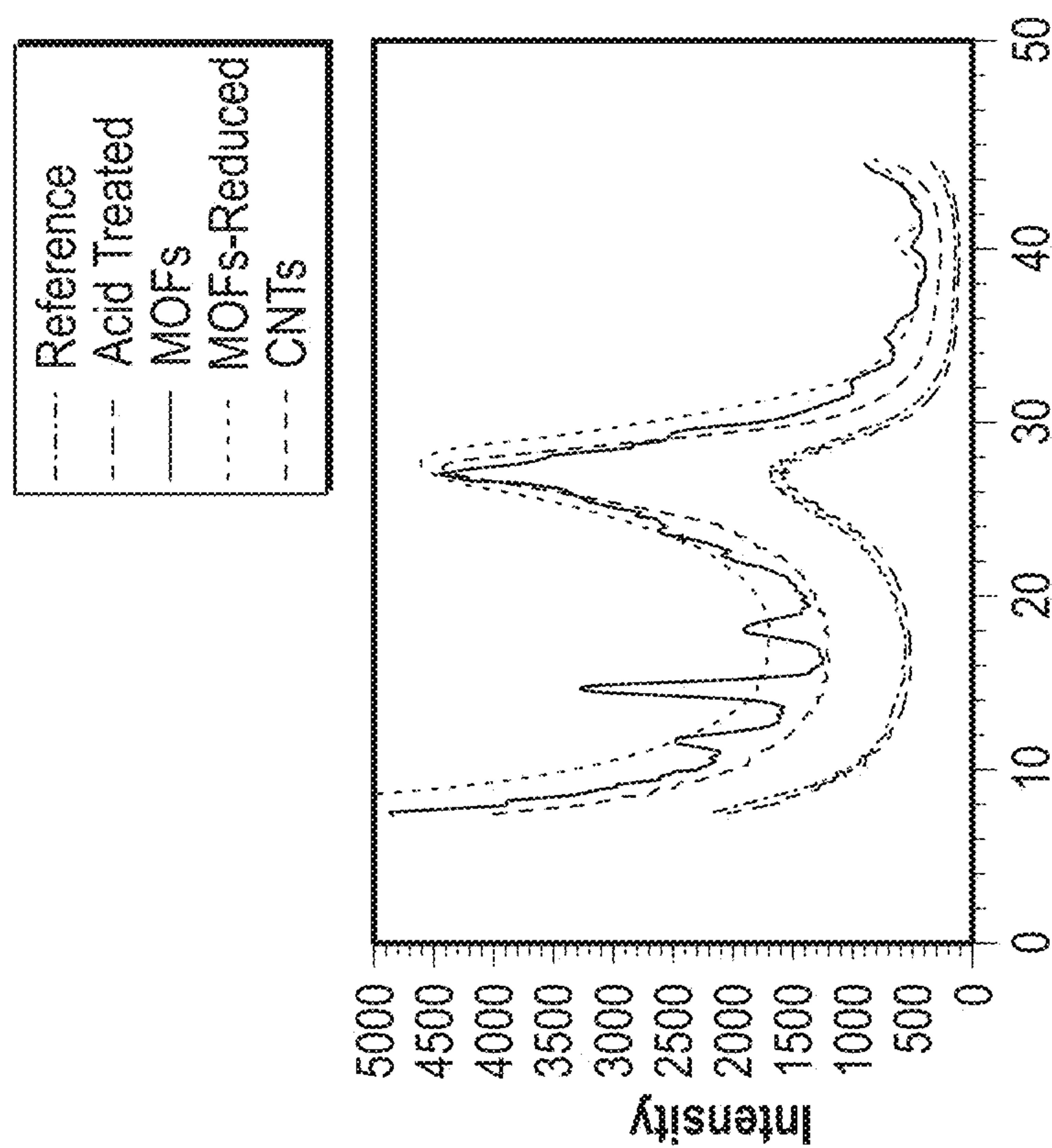


FIG. 8



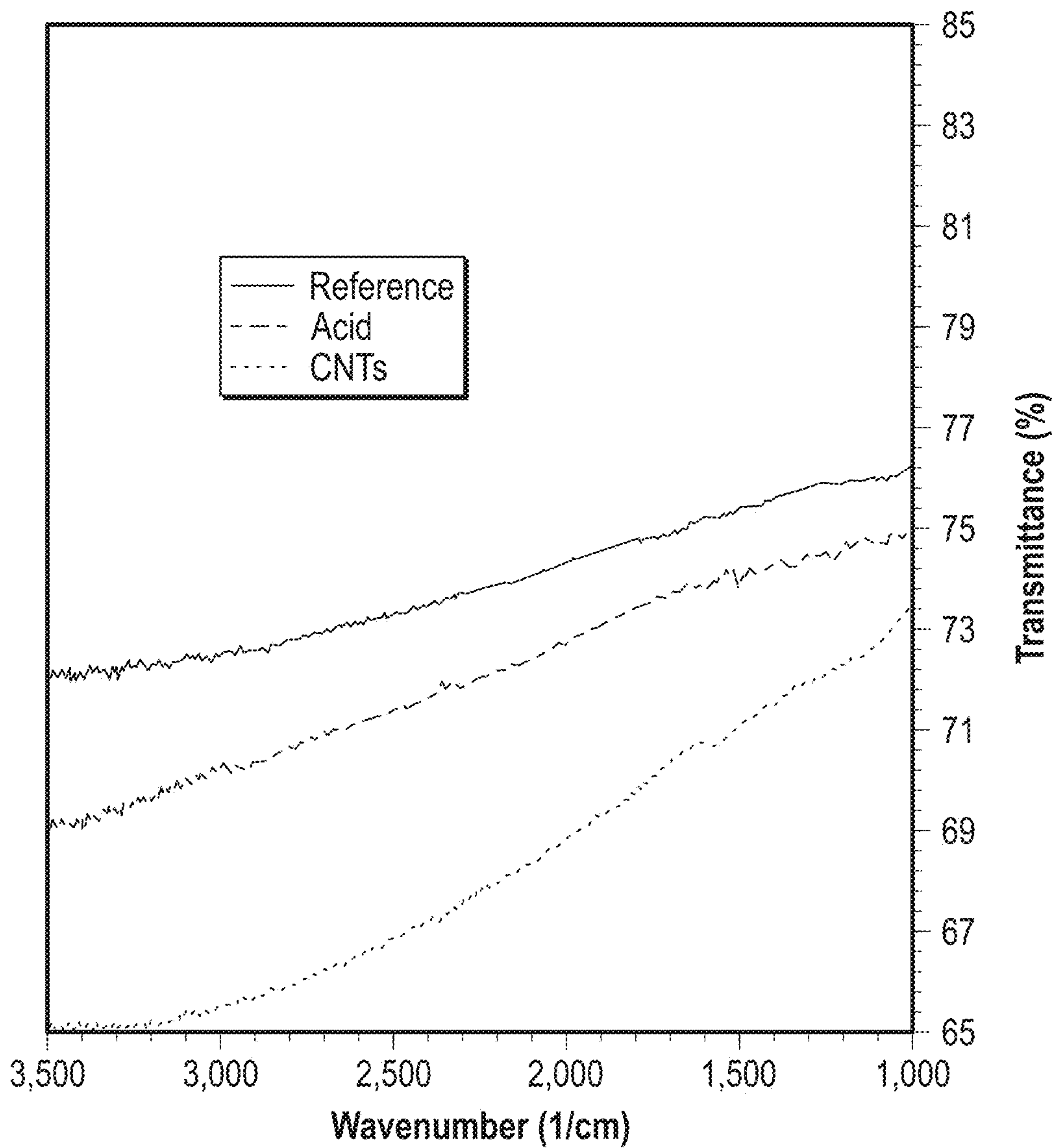


FIG. 10



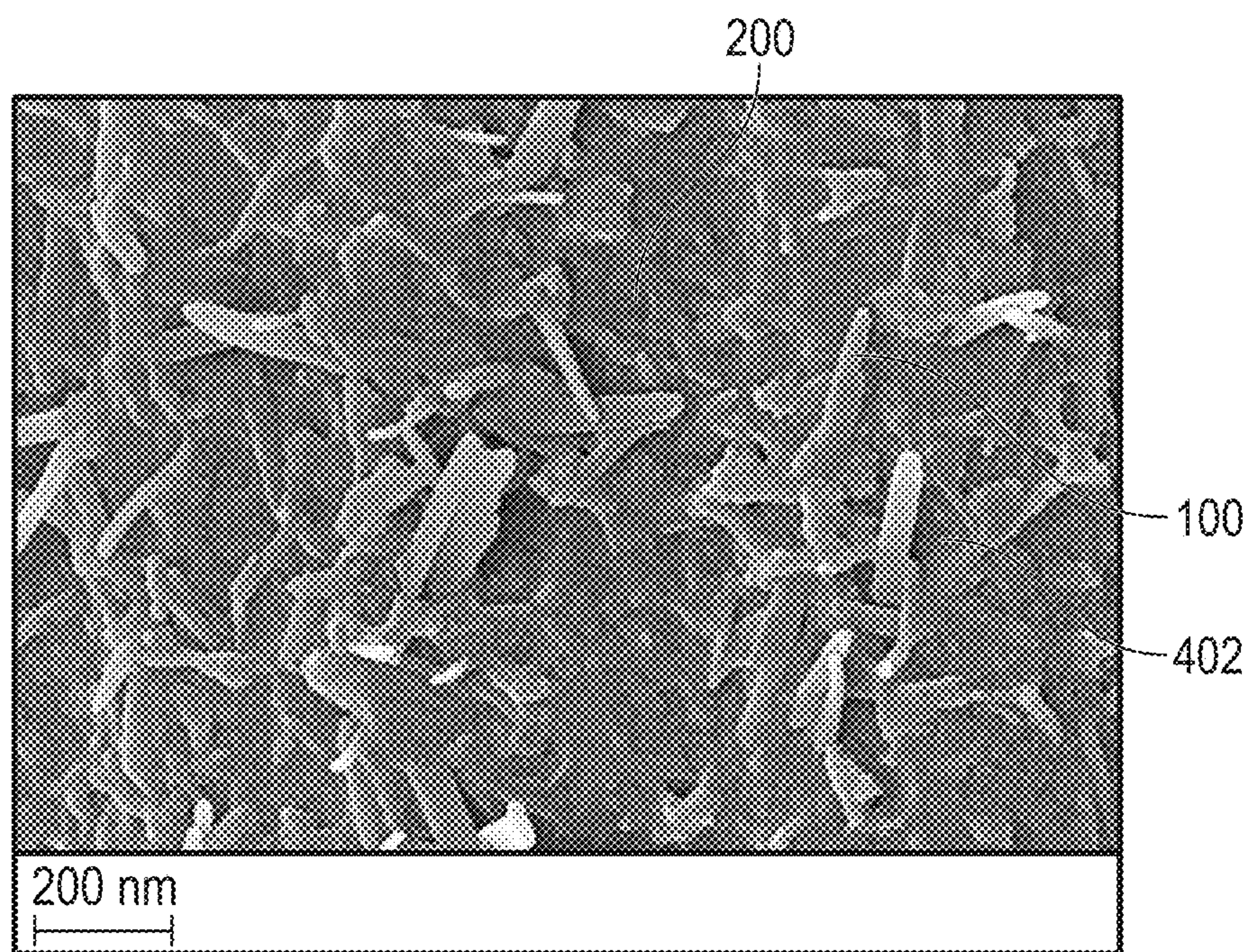


FIG. 11A

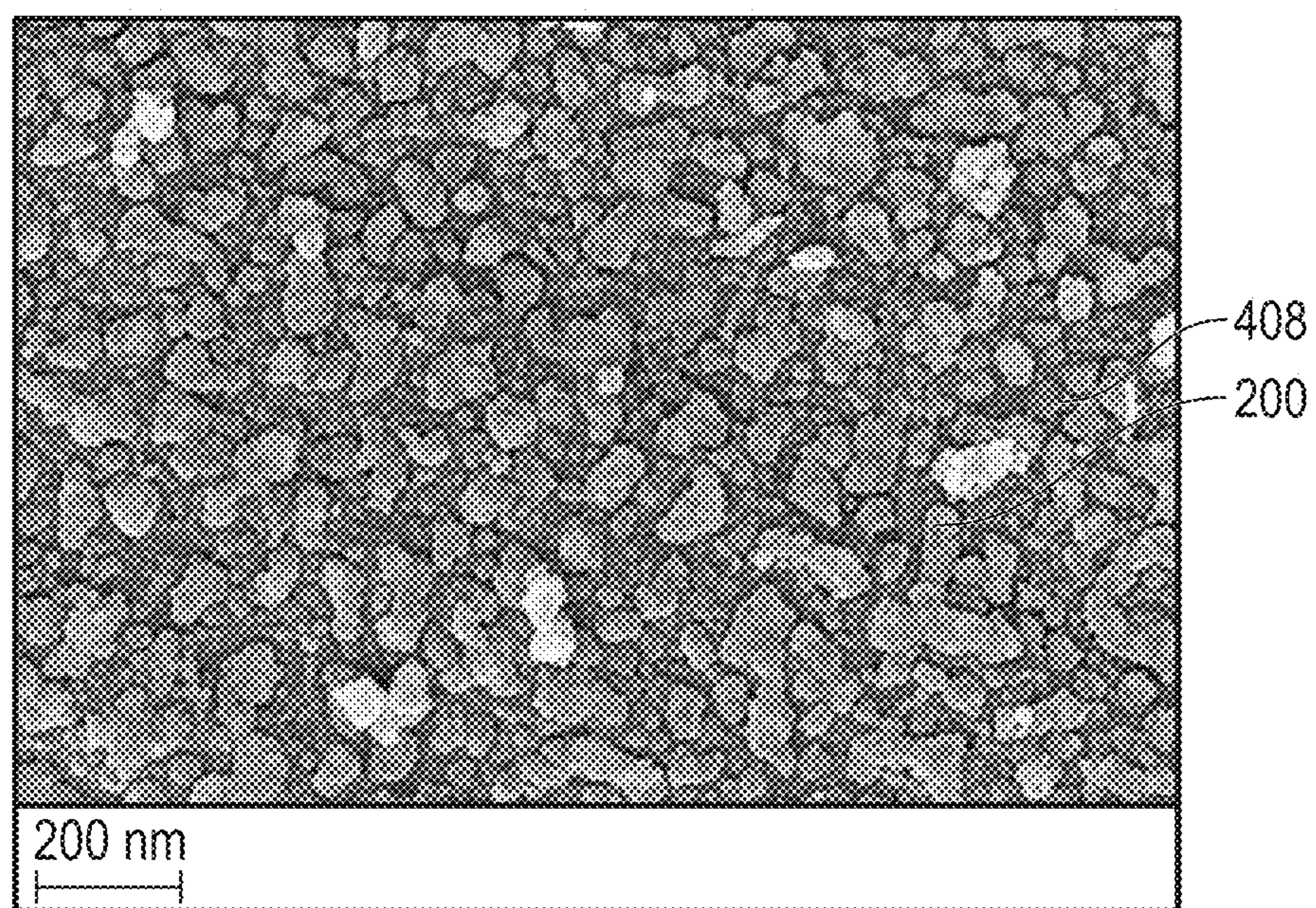


FIG. 11B



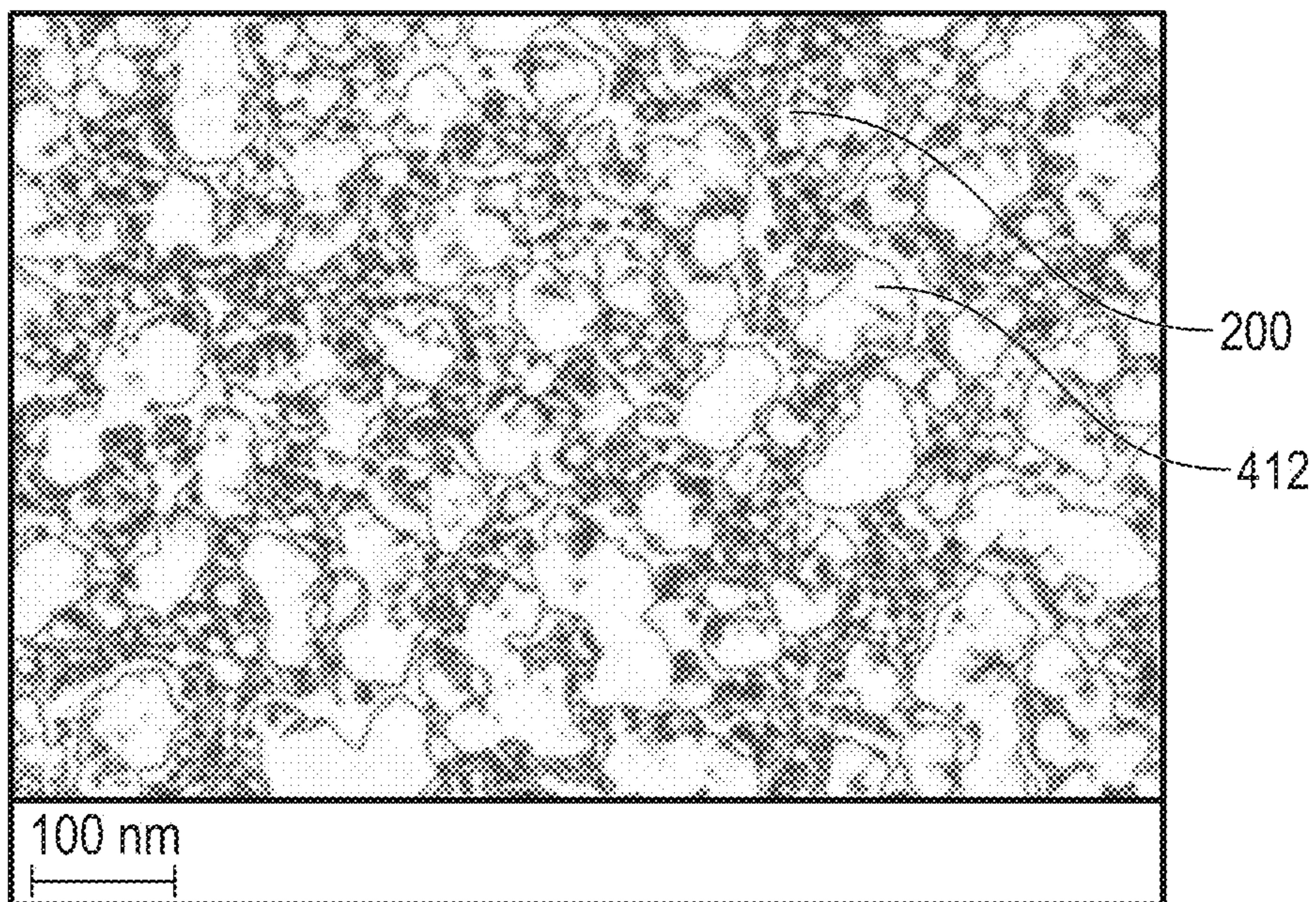


FIG. 12A

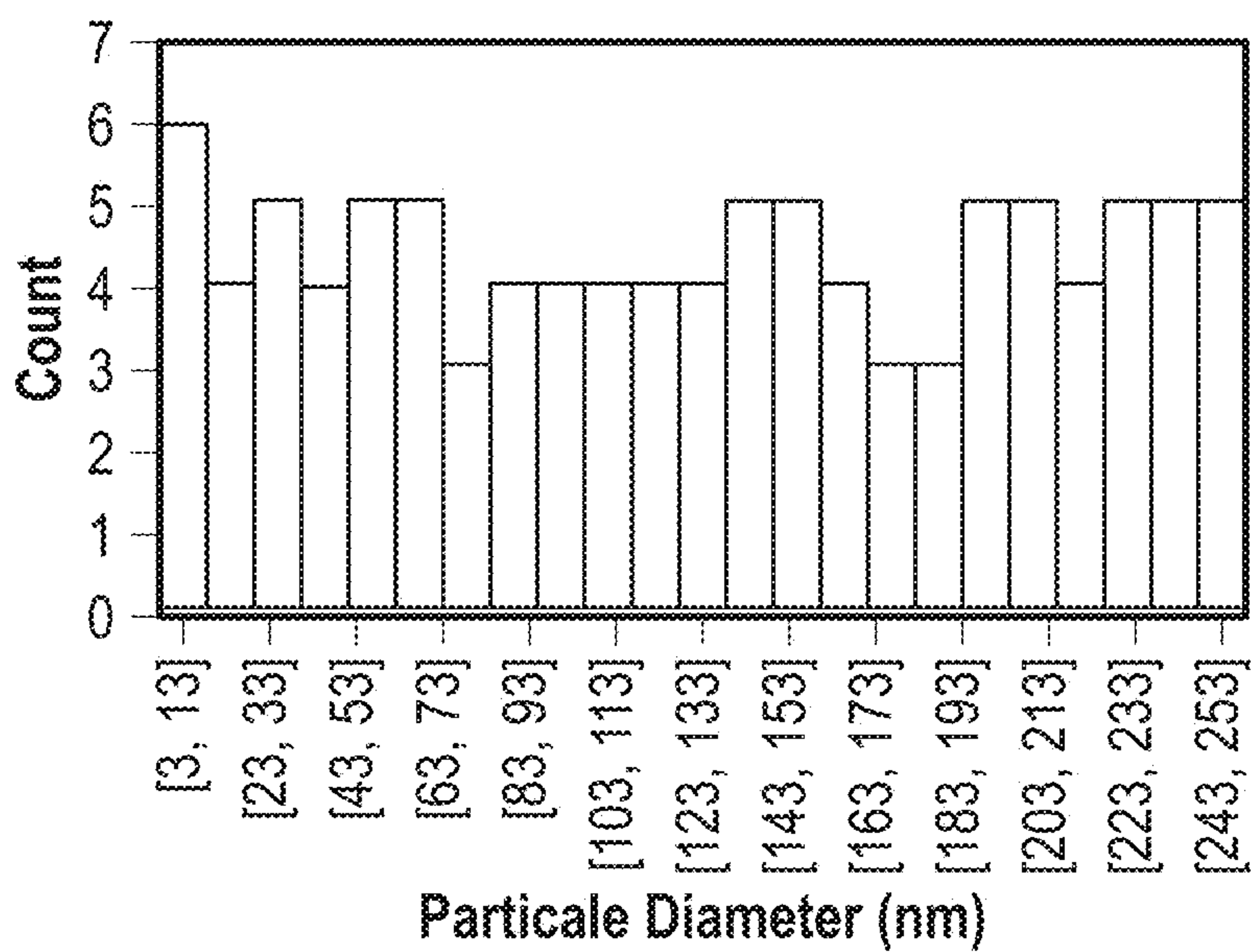


FIG. 12B



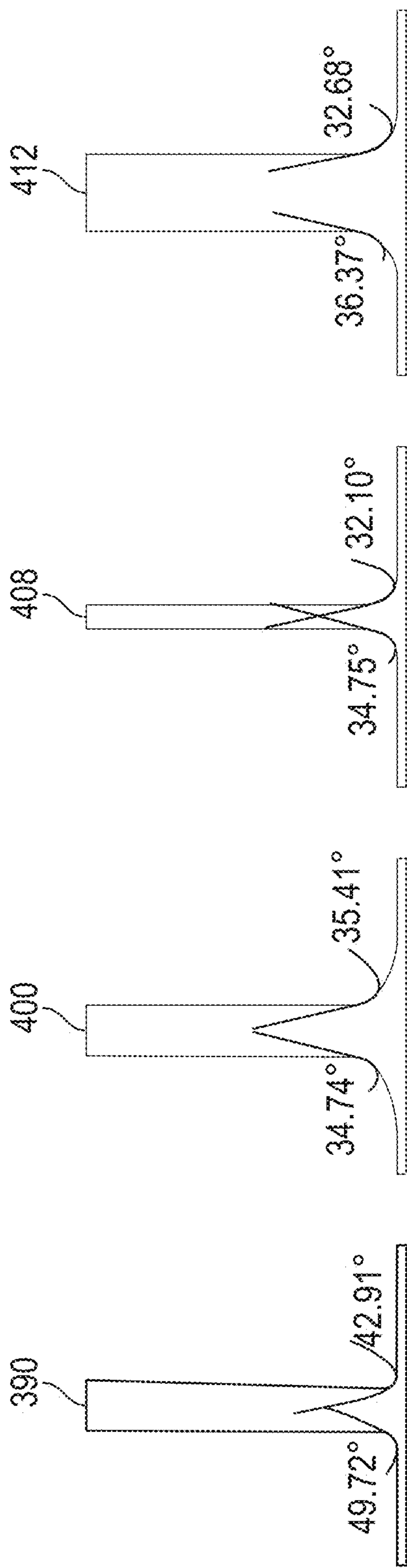


FIG. 13A

FIG. 13B

FIG. 13C

FIG. 13D



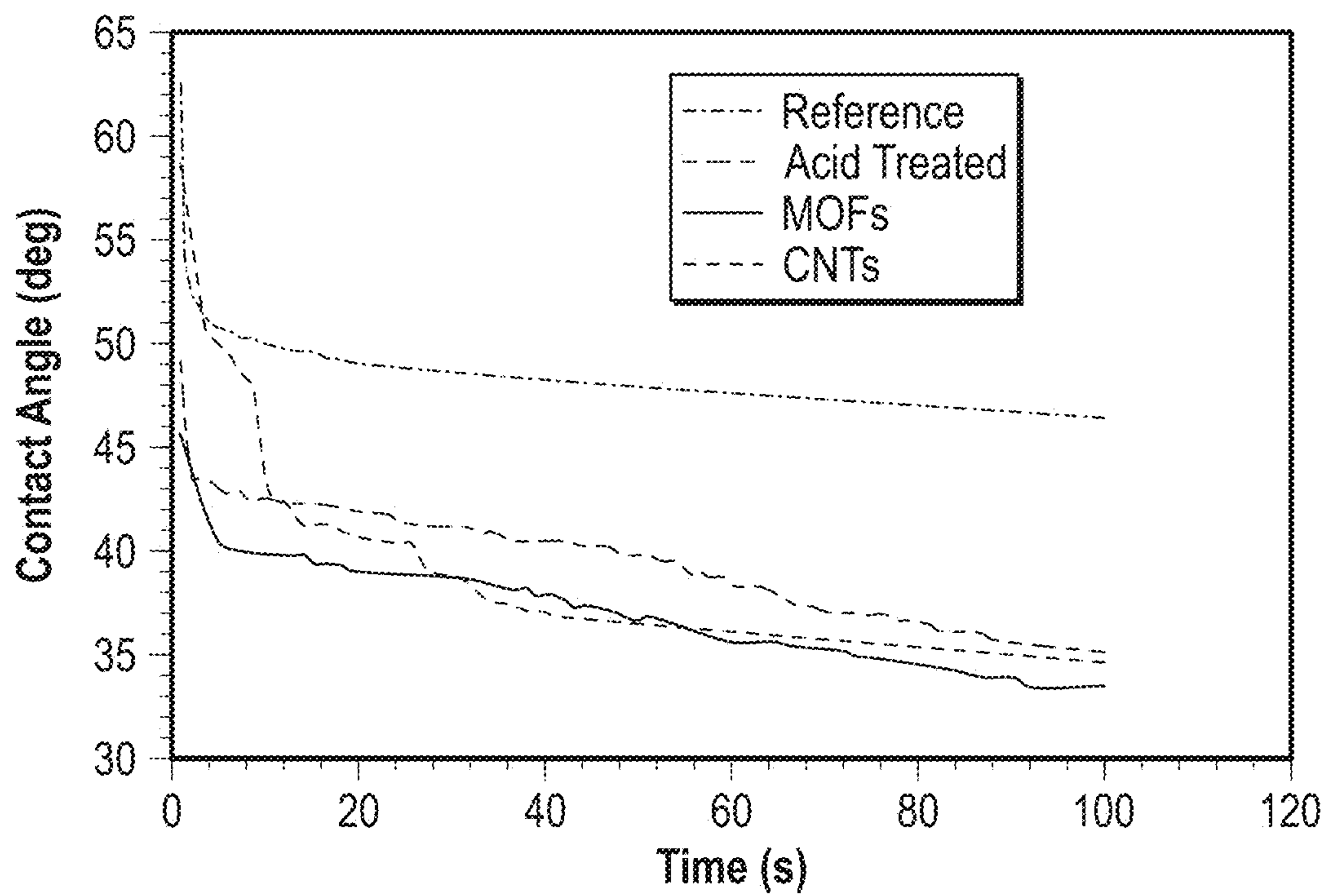


FIG. 14

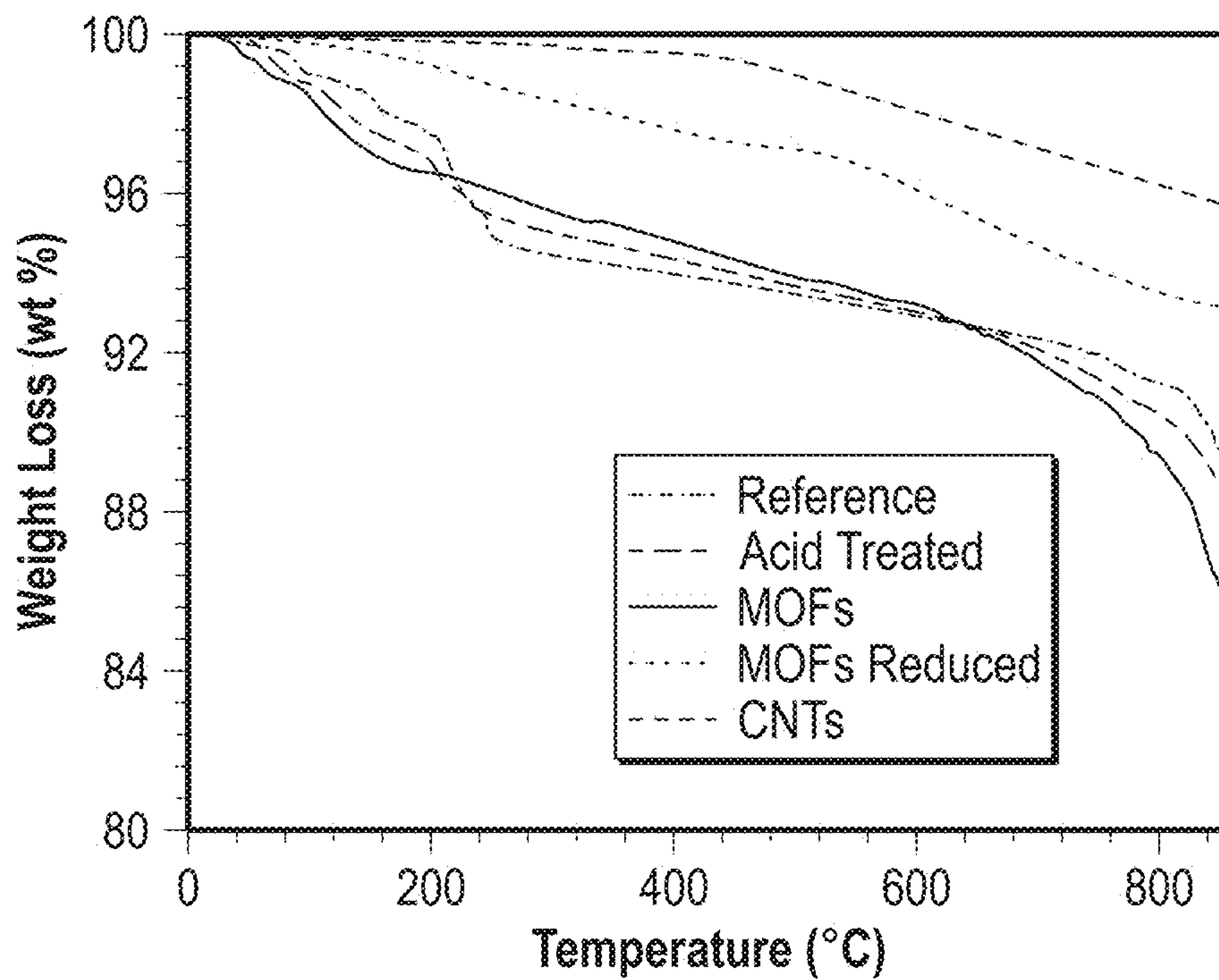


FIG. 15



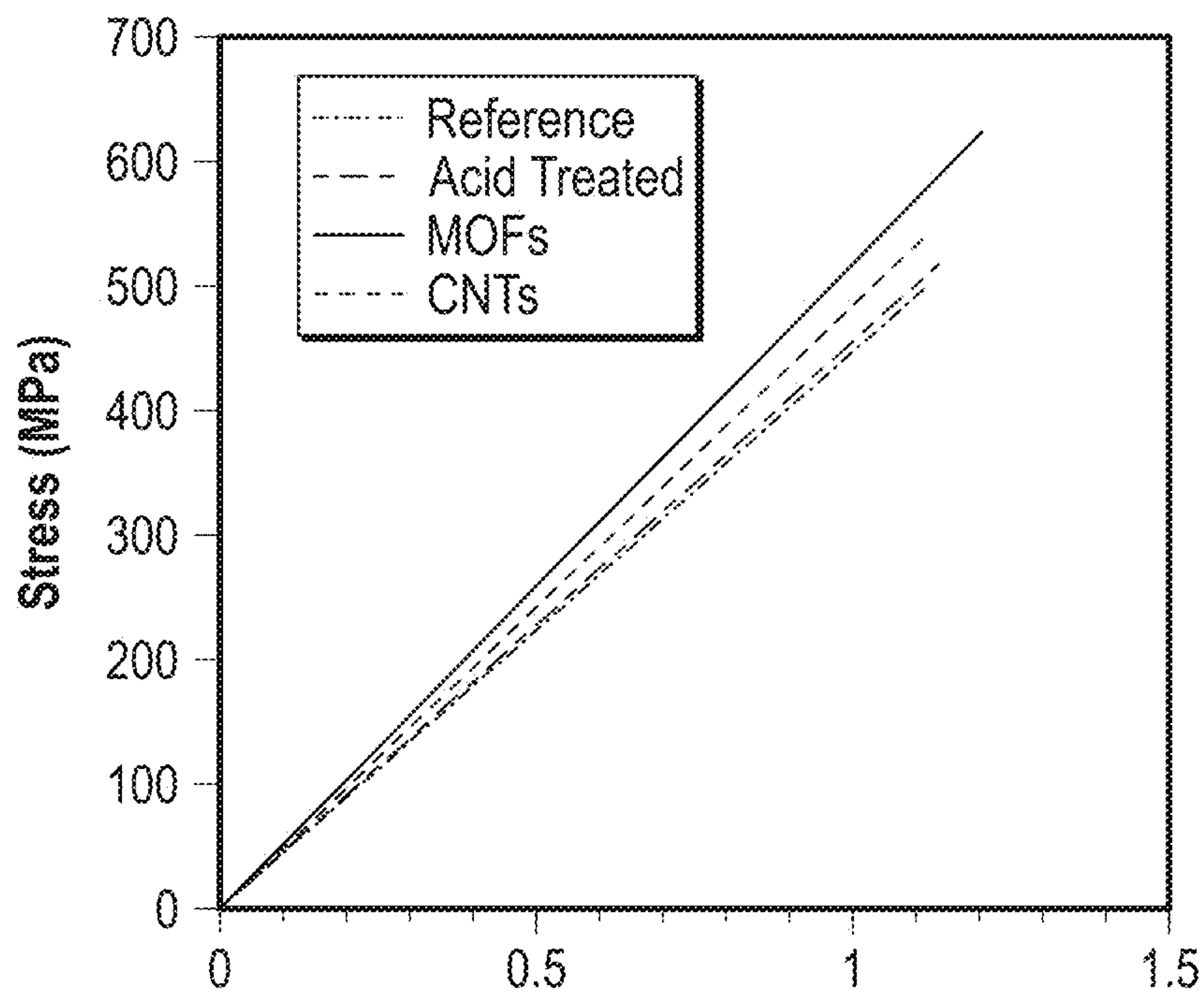


FIG. 16A

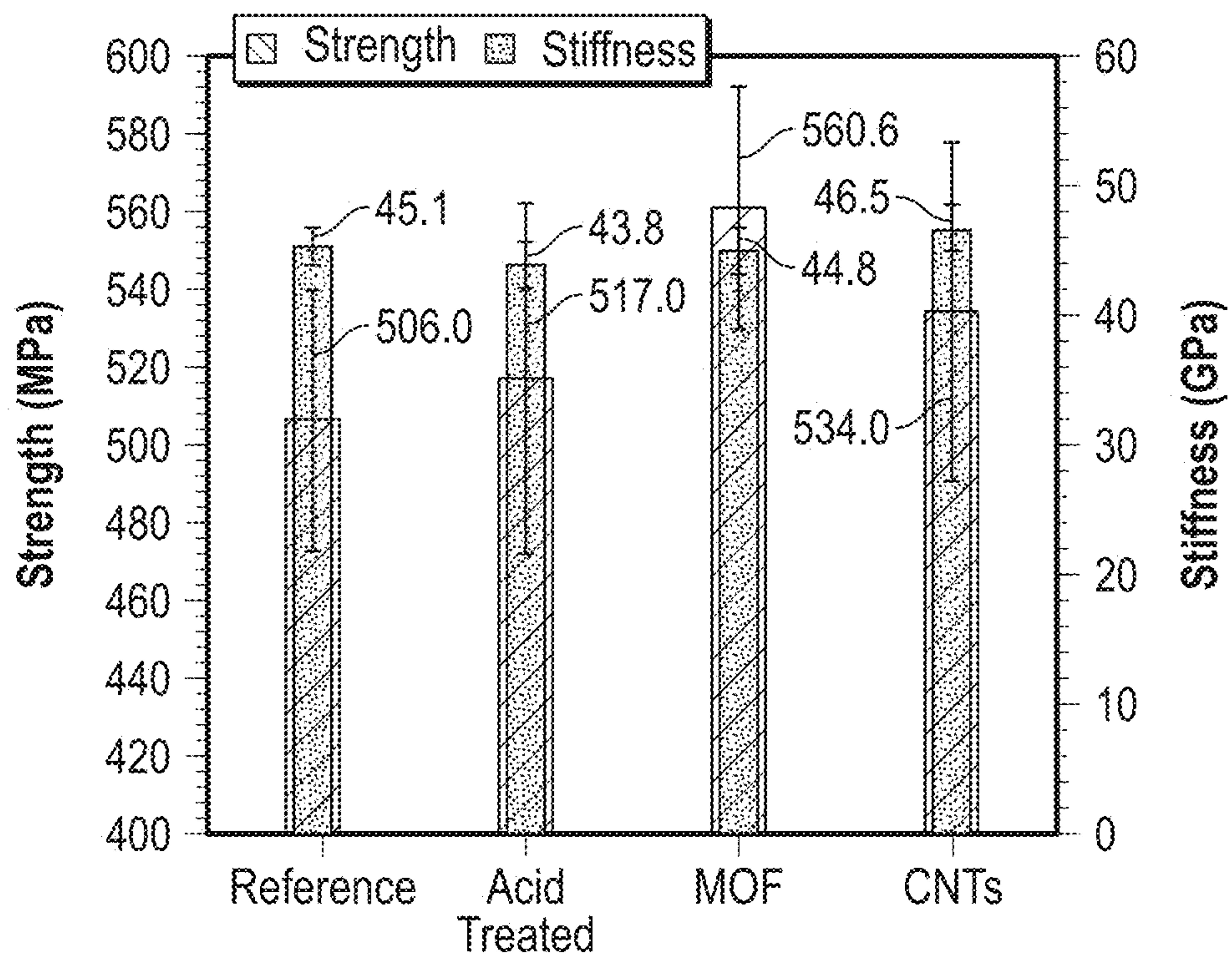


FIG. 16B



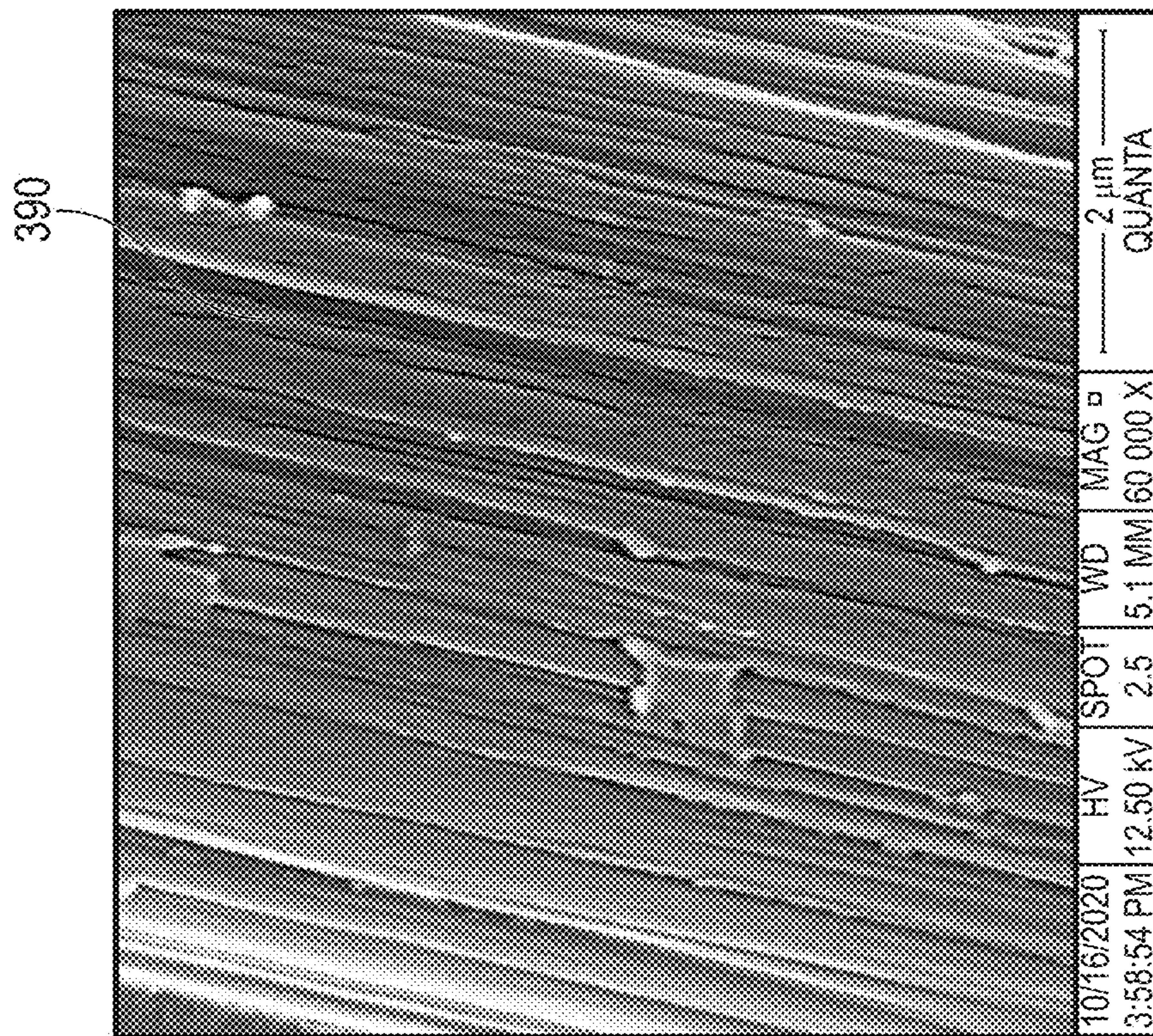


FIG. 17B

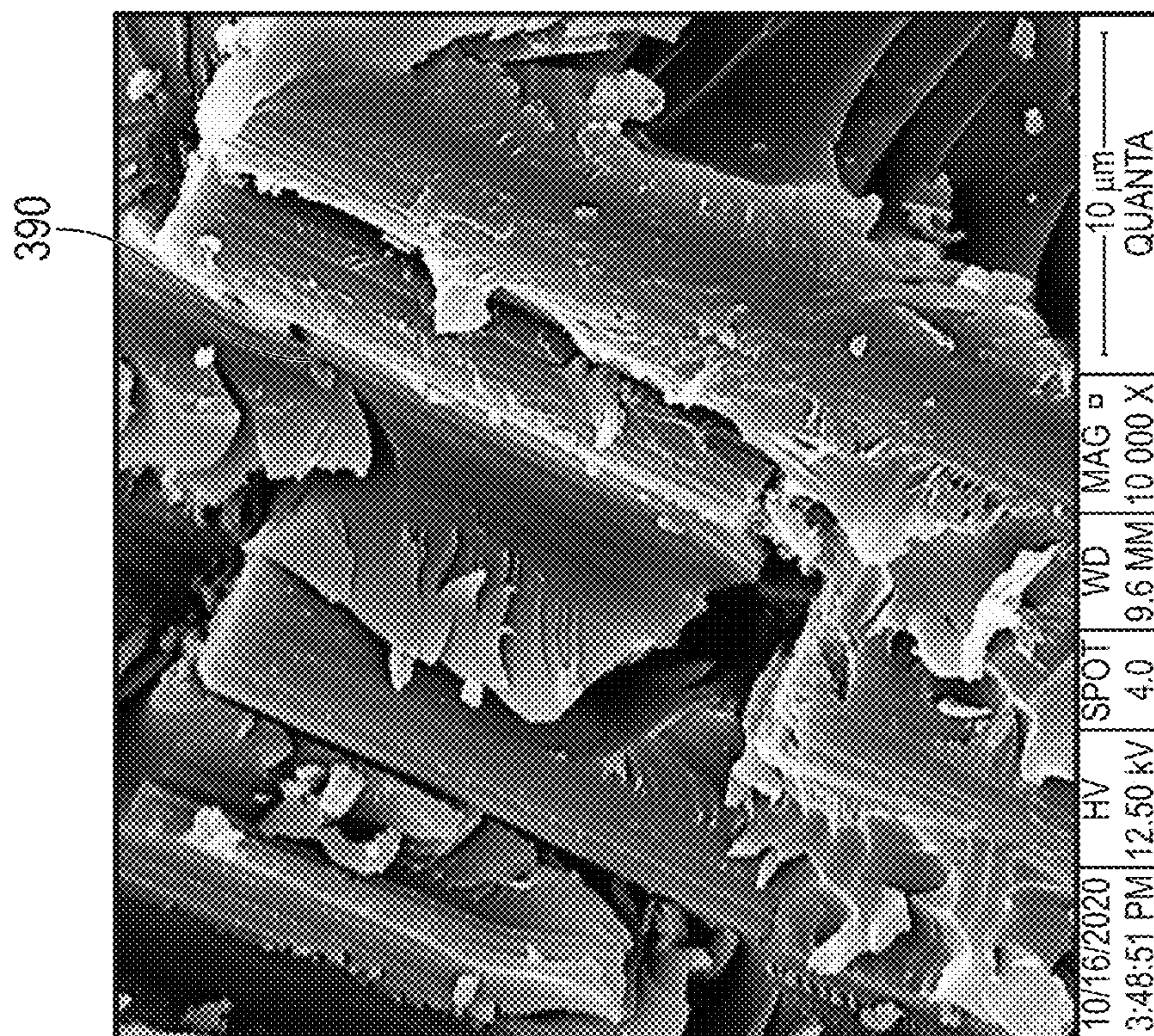


FIG. 17A



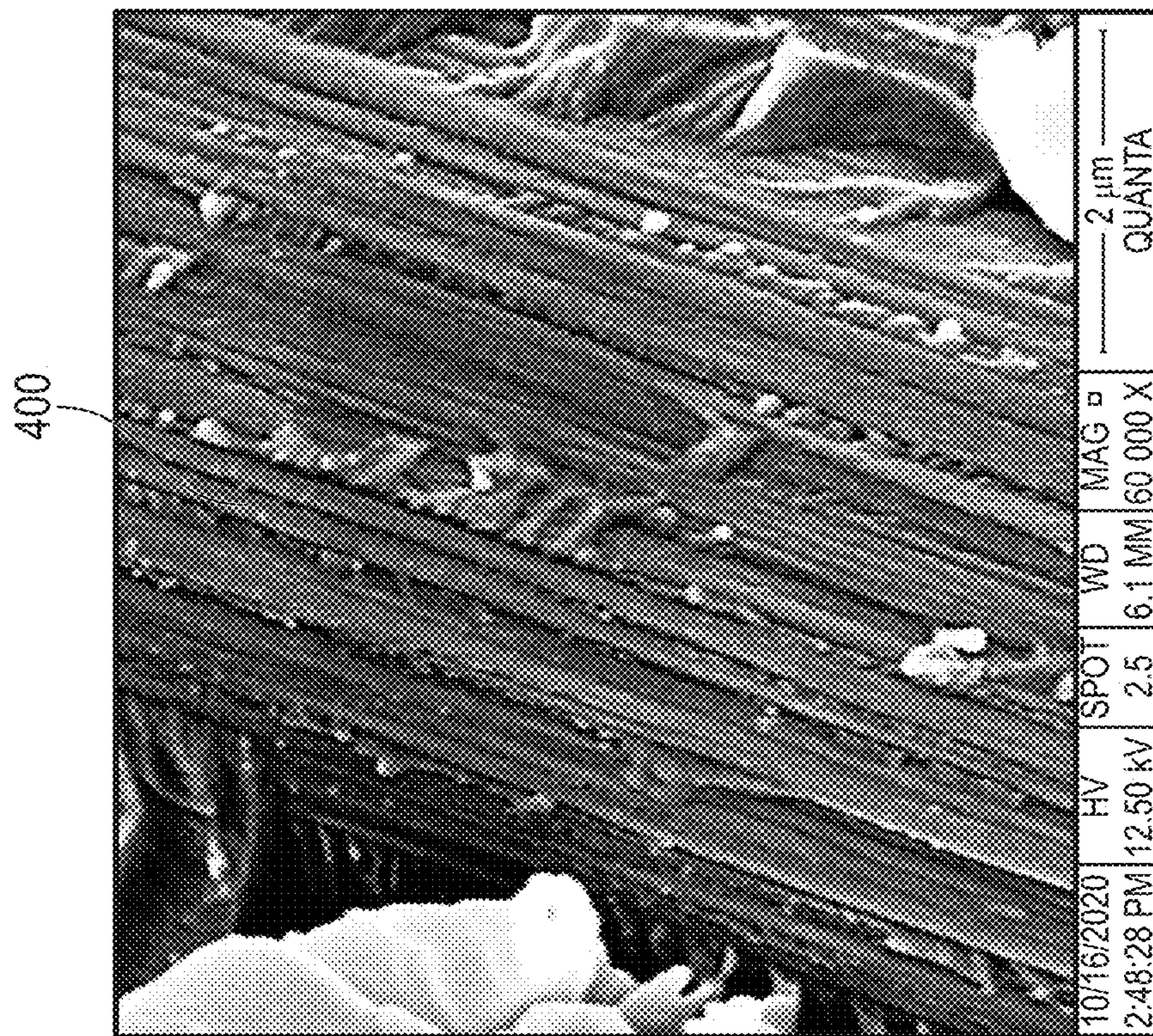


FIG. 17D

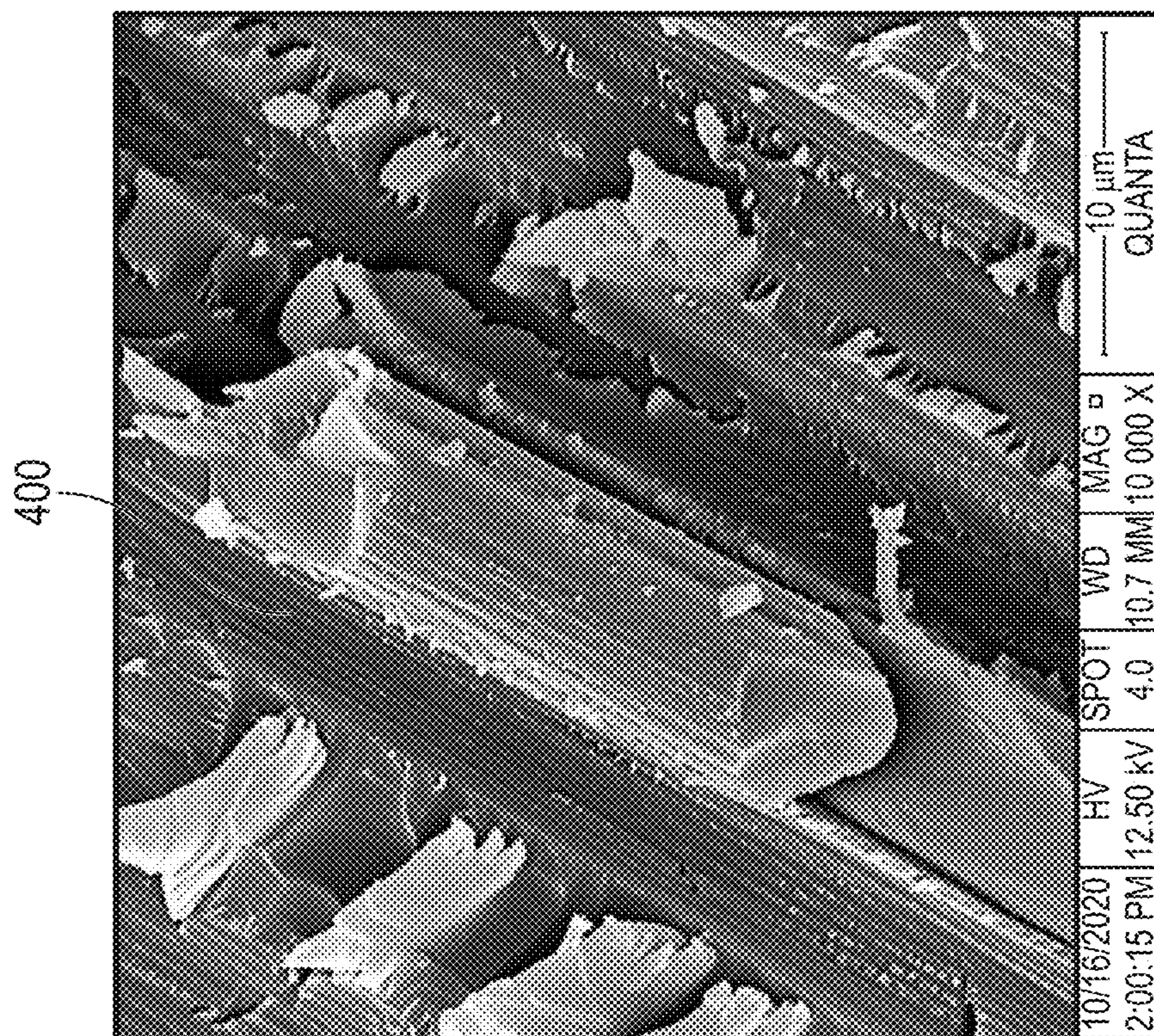


FIG. 17C



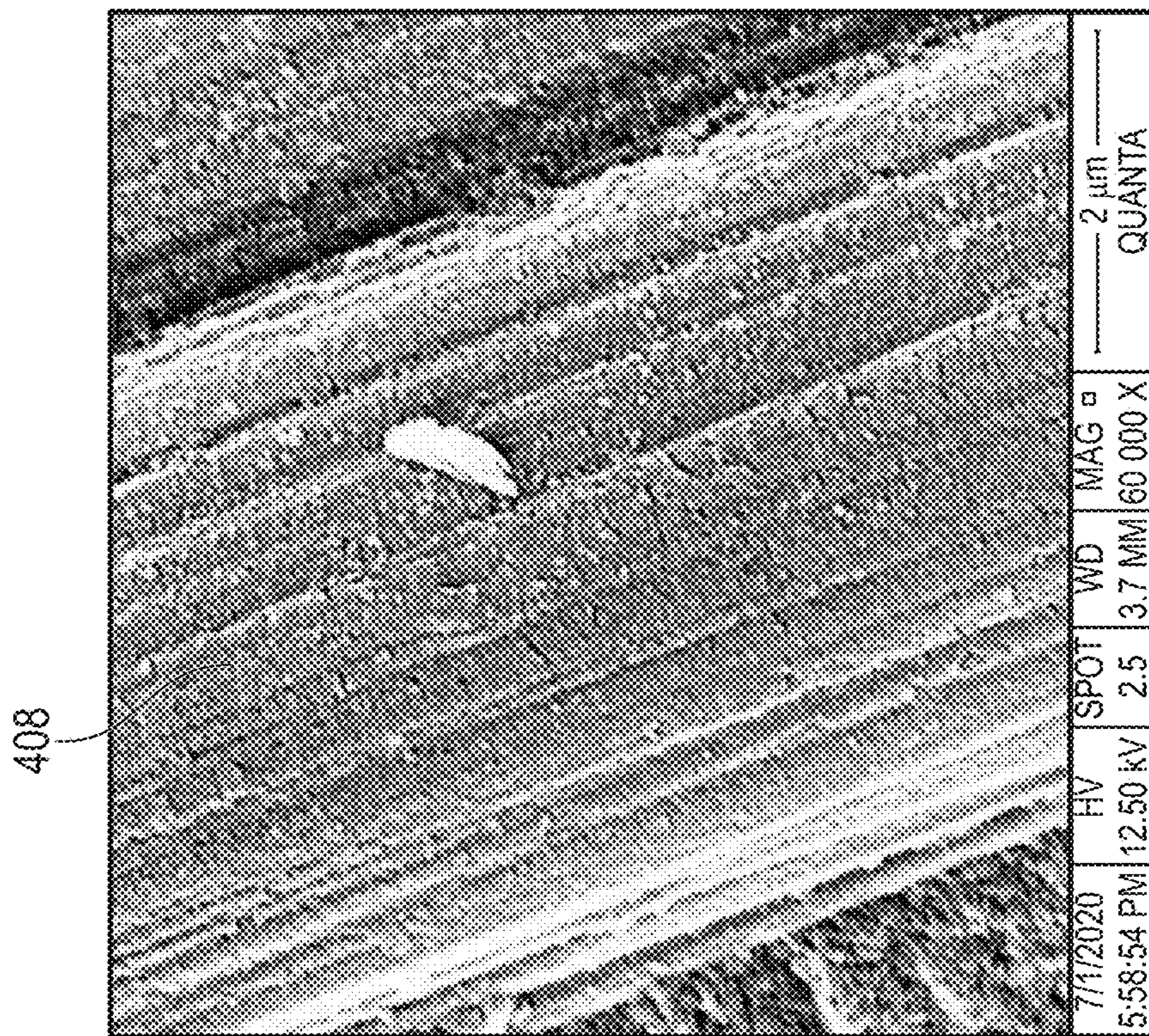


FIG. 18B

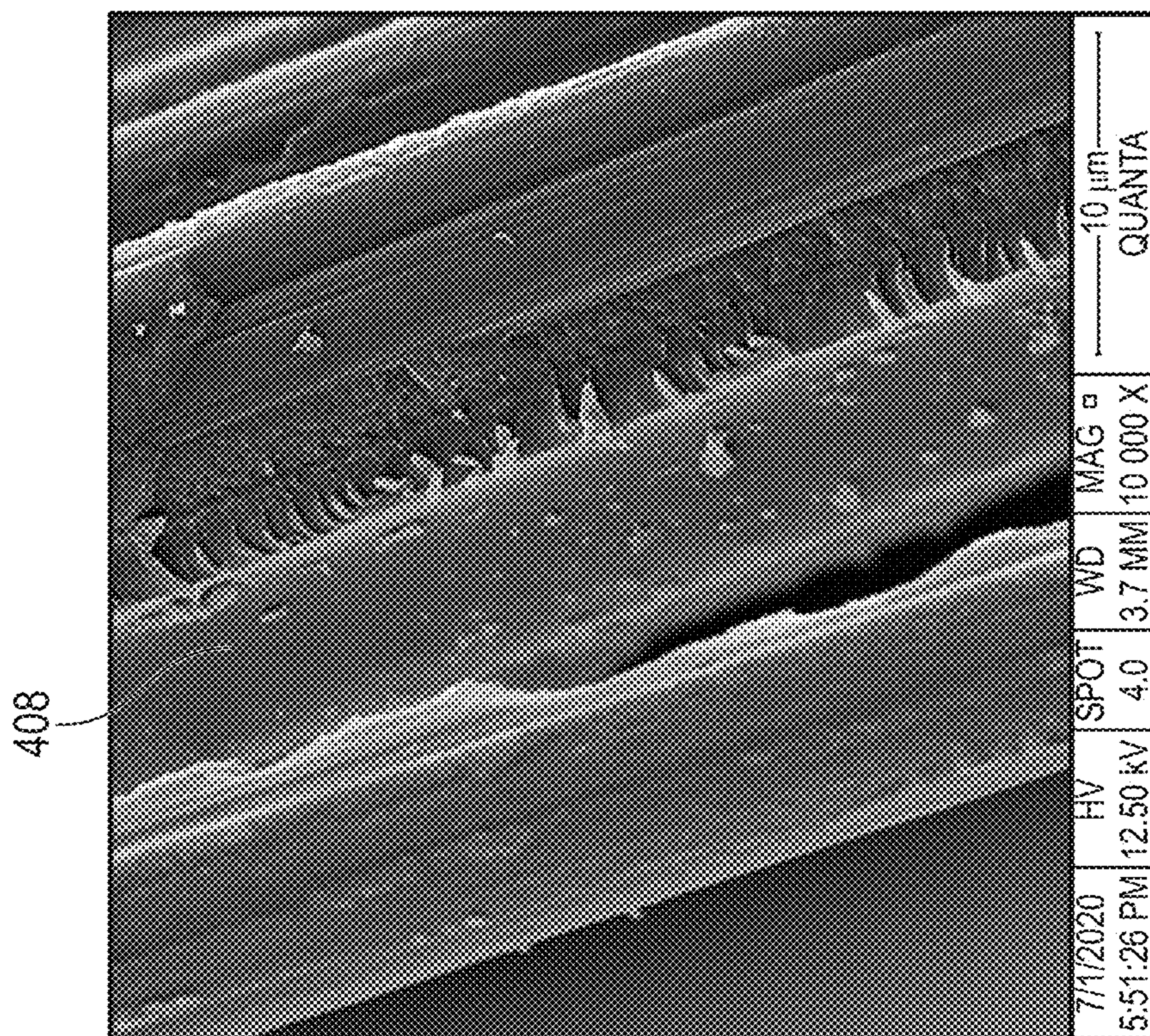


FIG. 18A



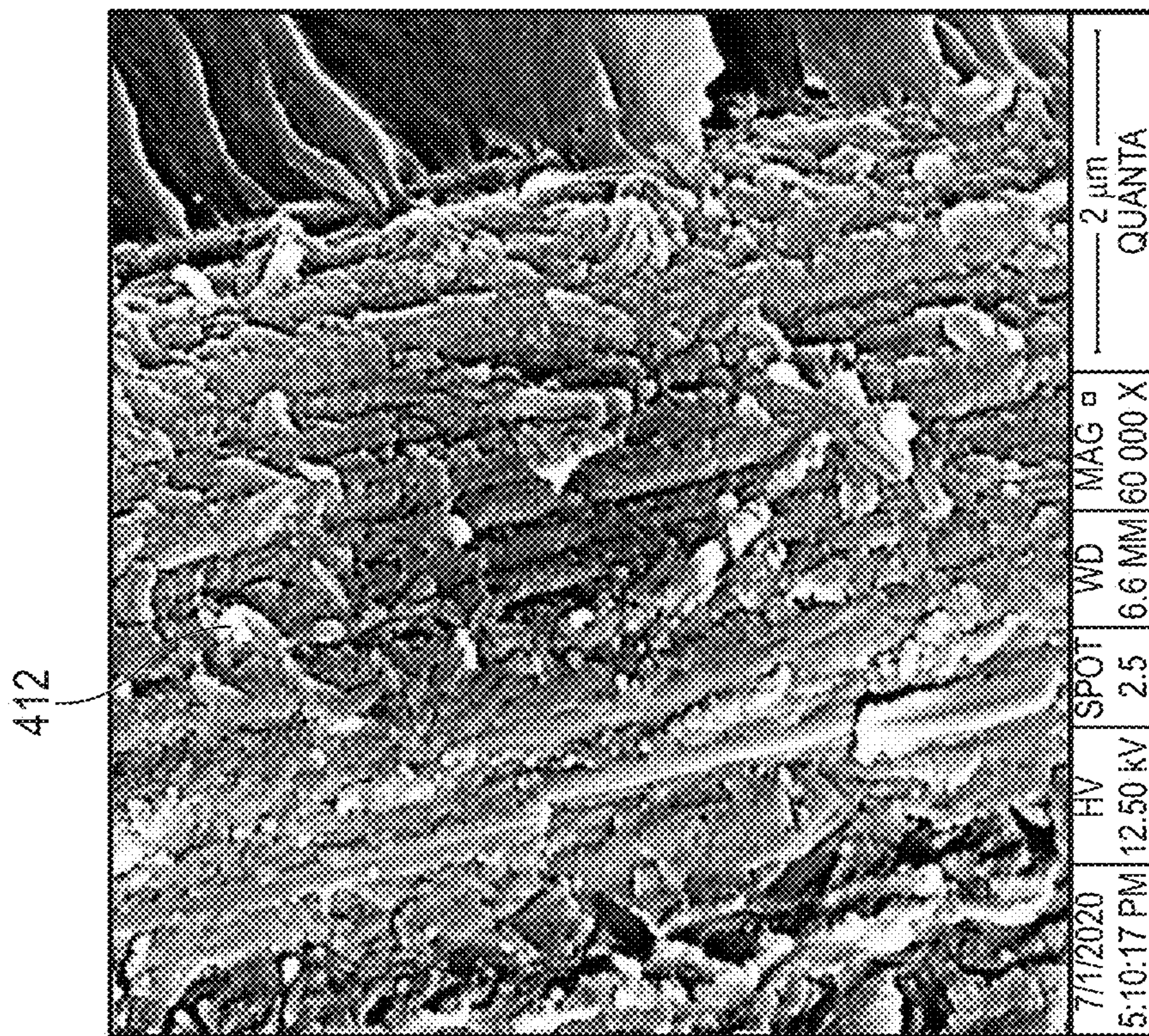


FIG. 18D

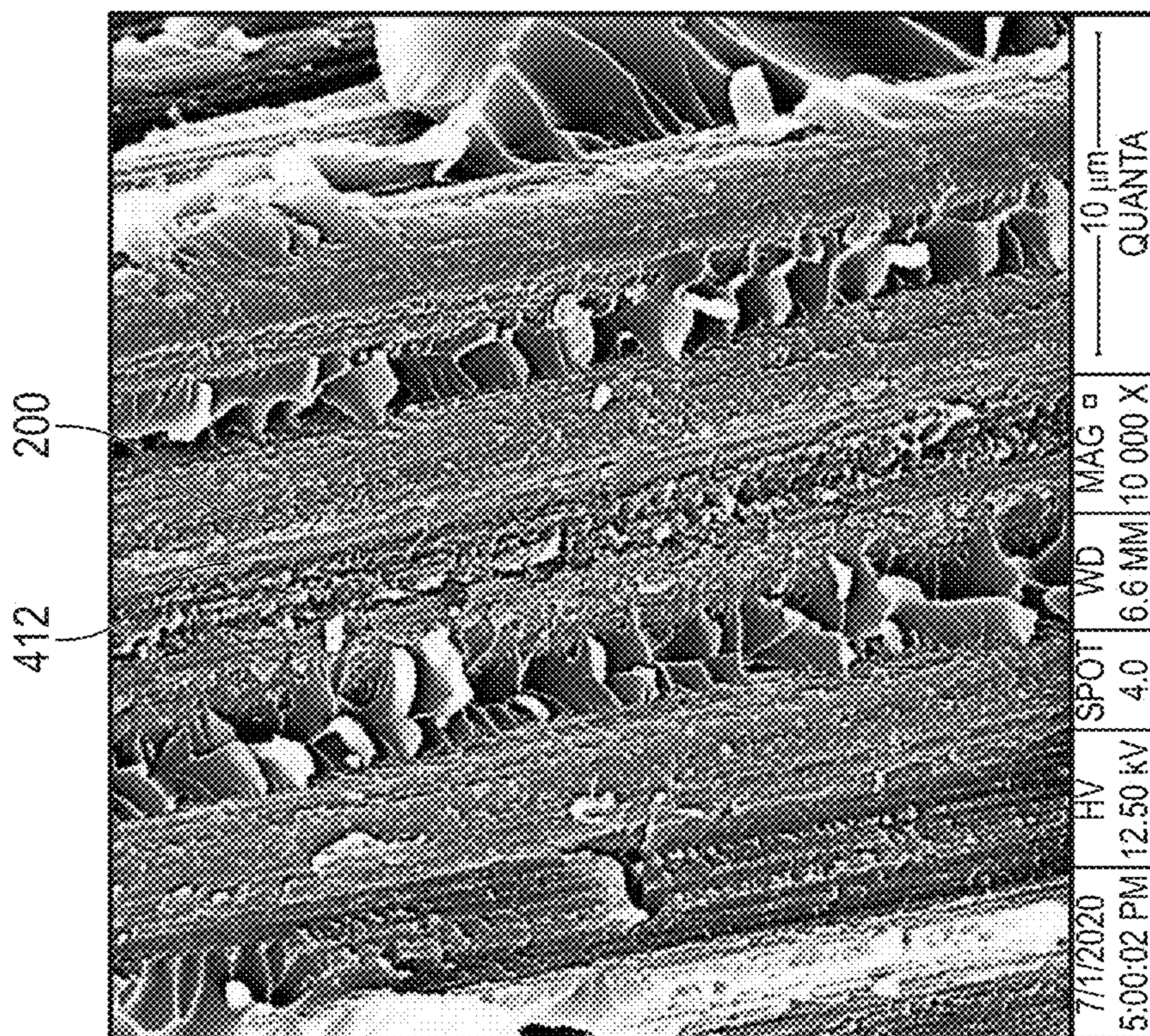


FIG. 18C



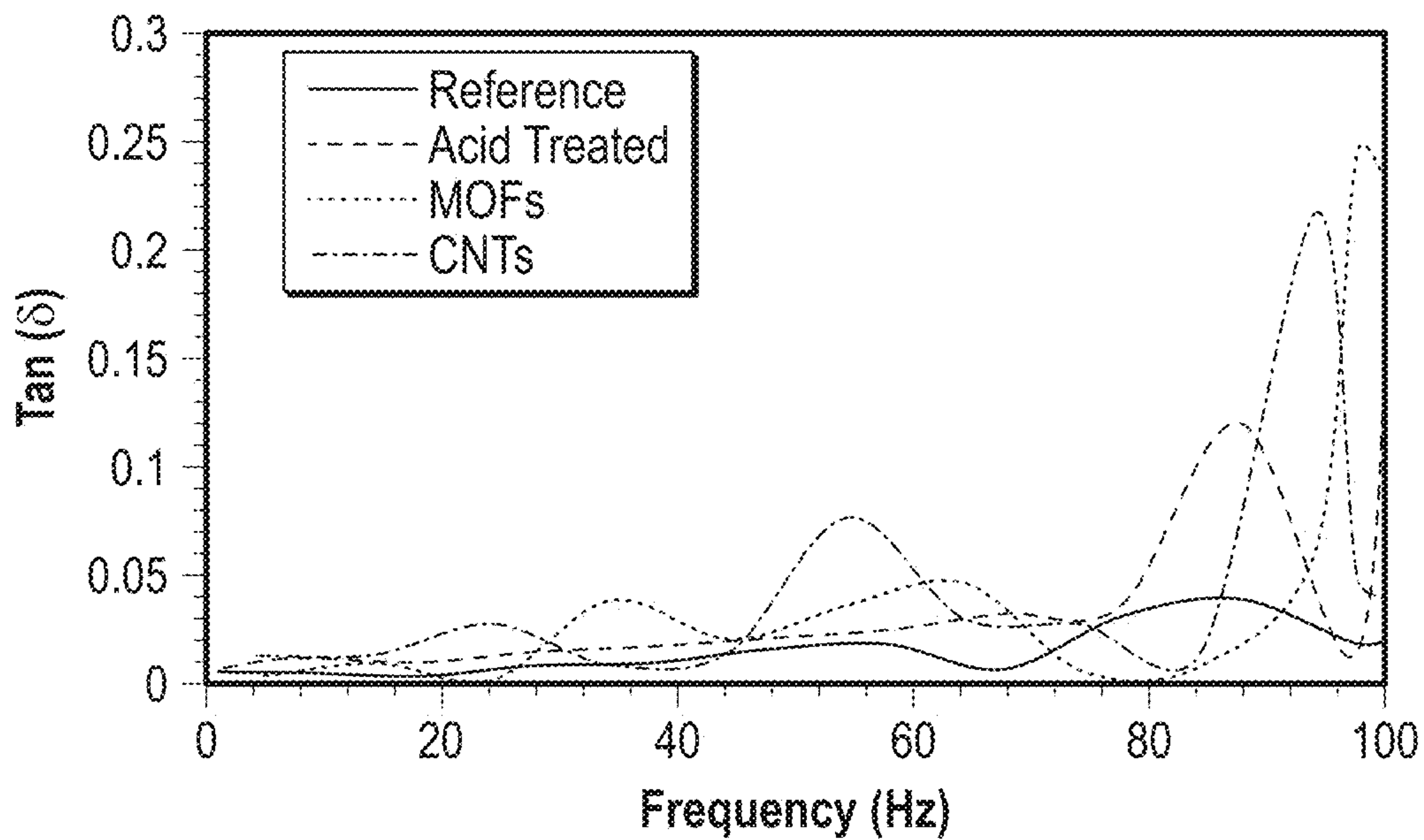


FIG. 19

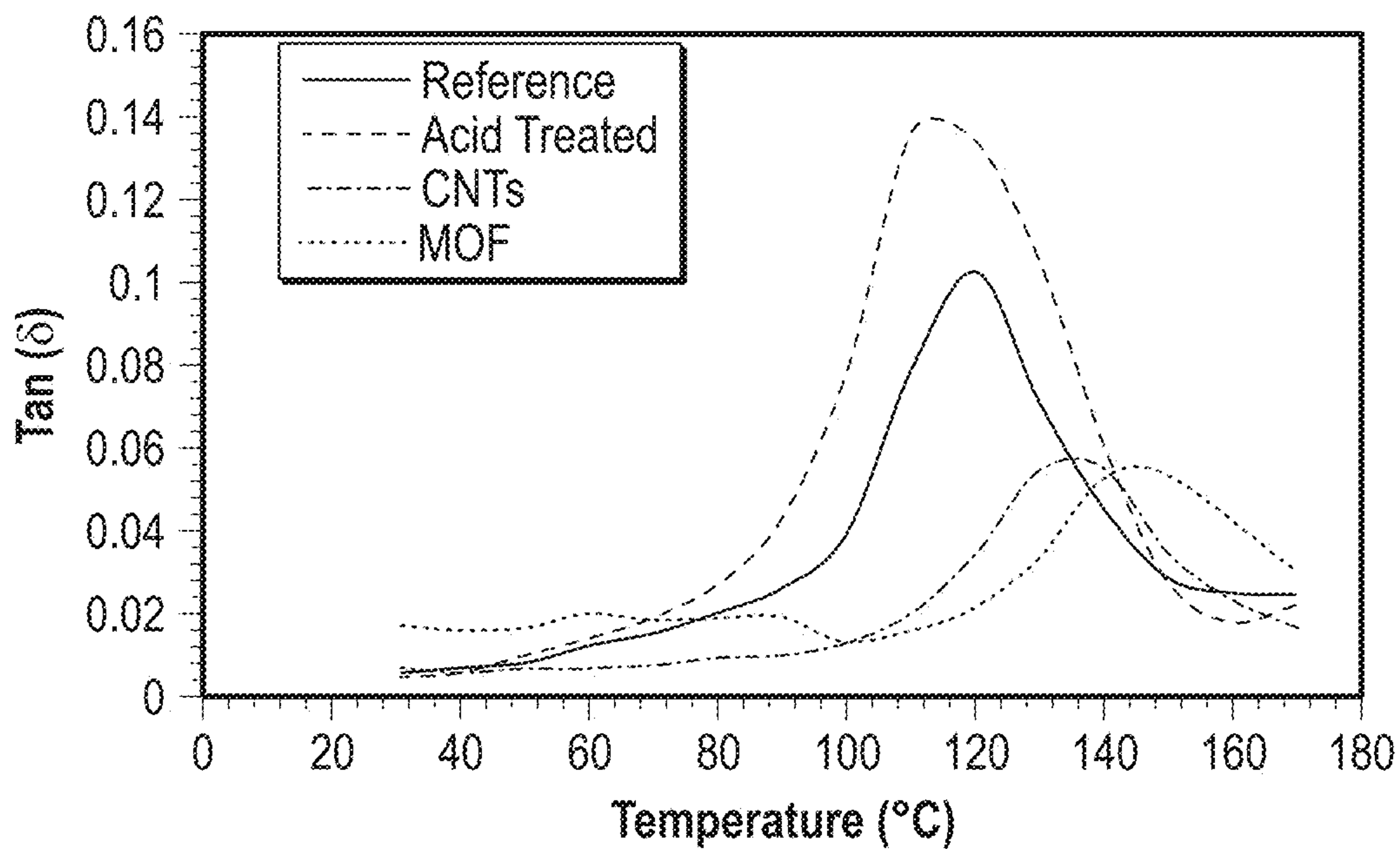


FIG. 20



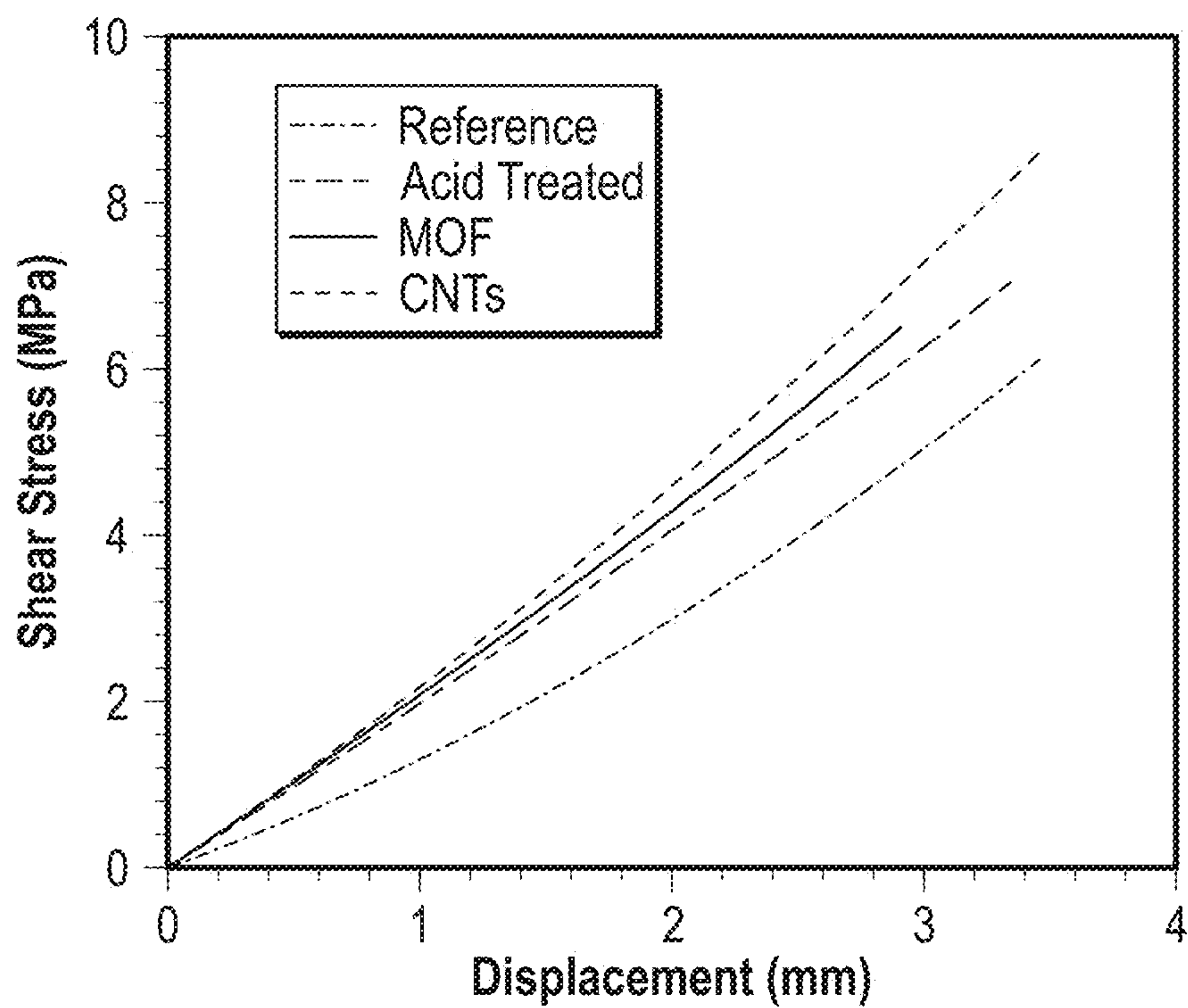


FIG. 21

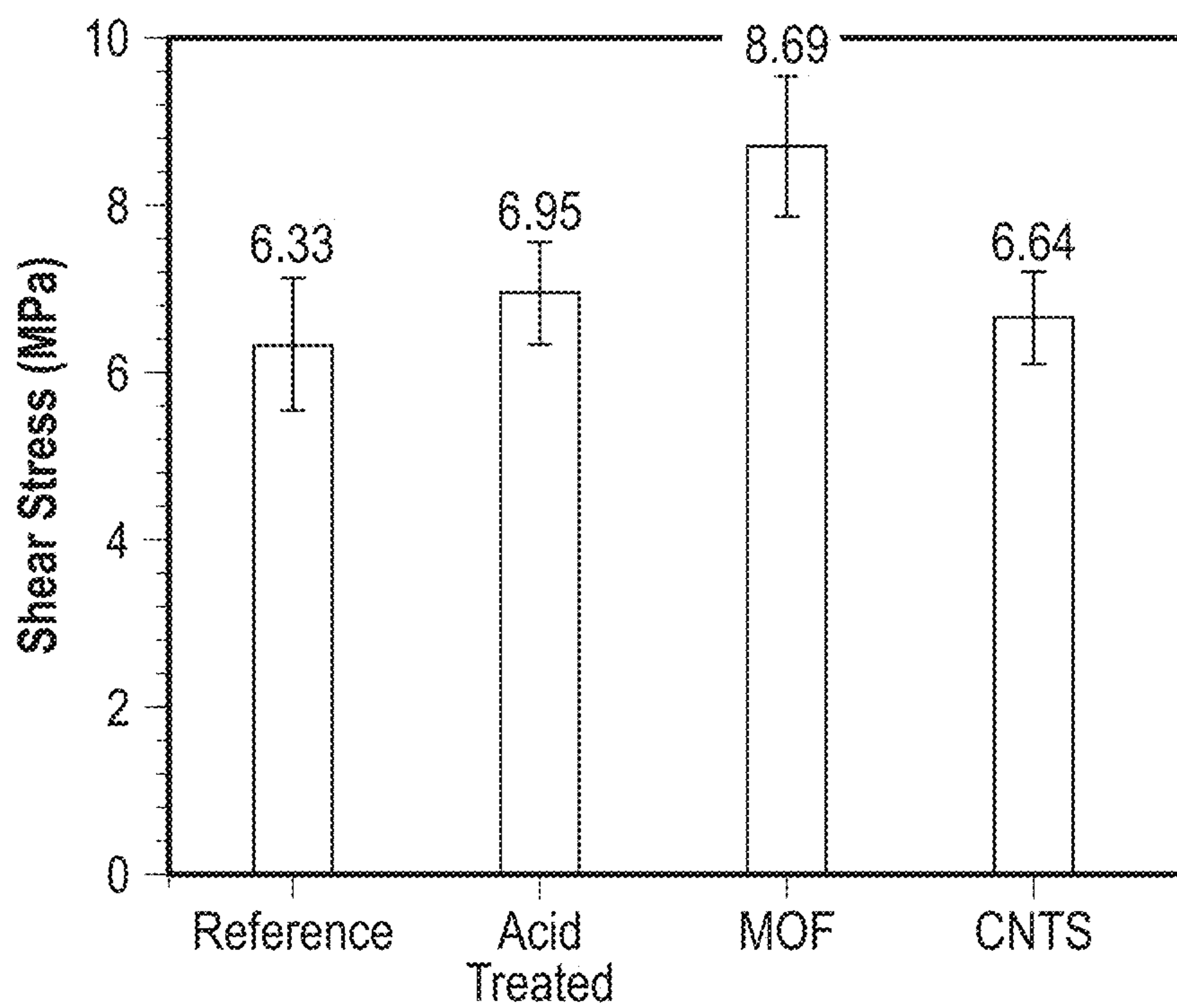


FIG. 22



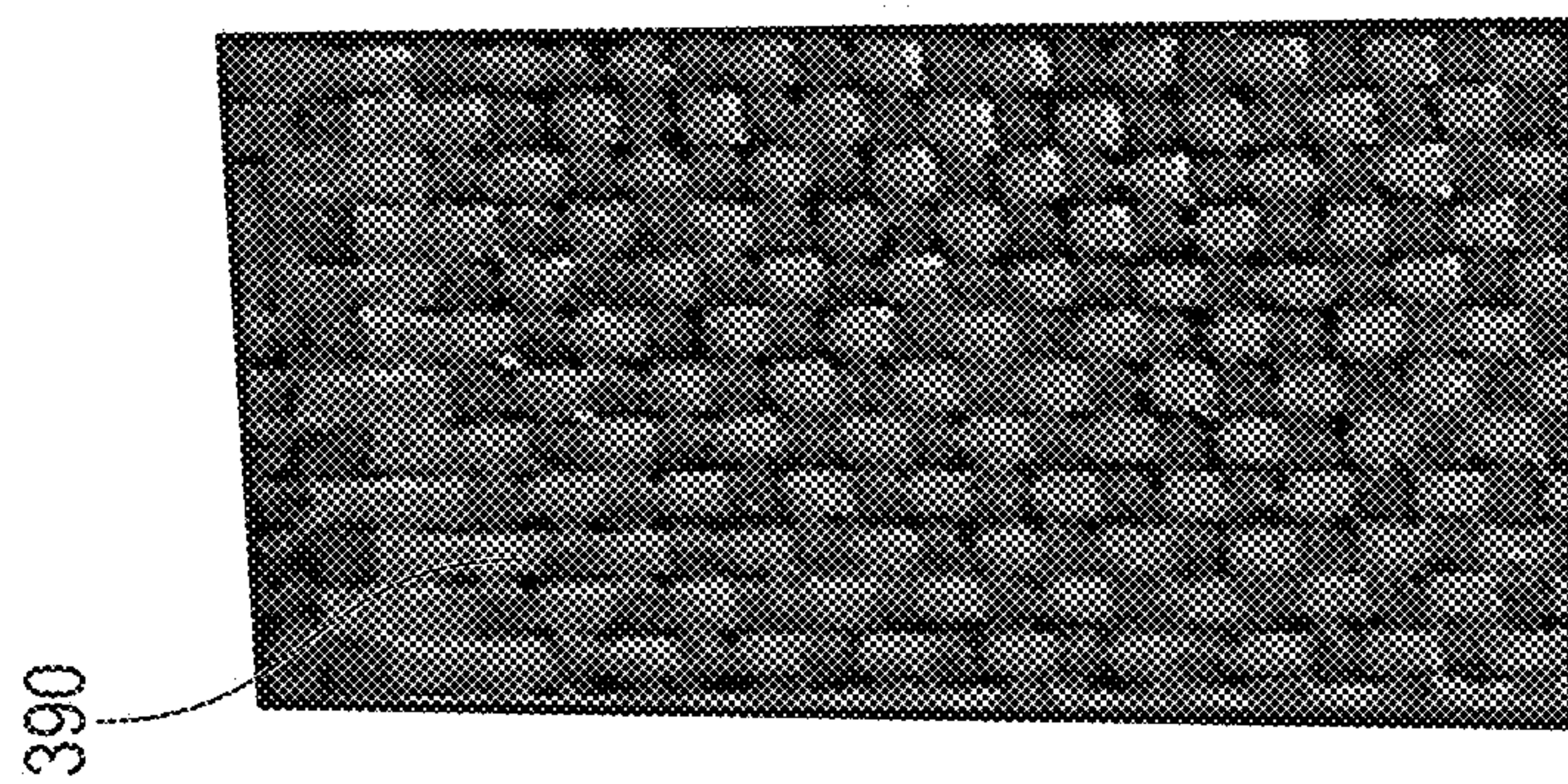


FIG. 23A

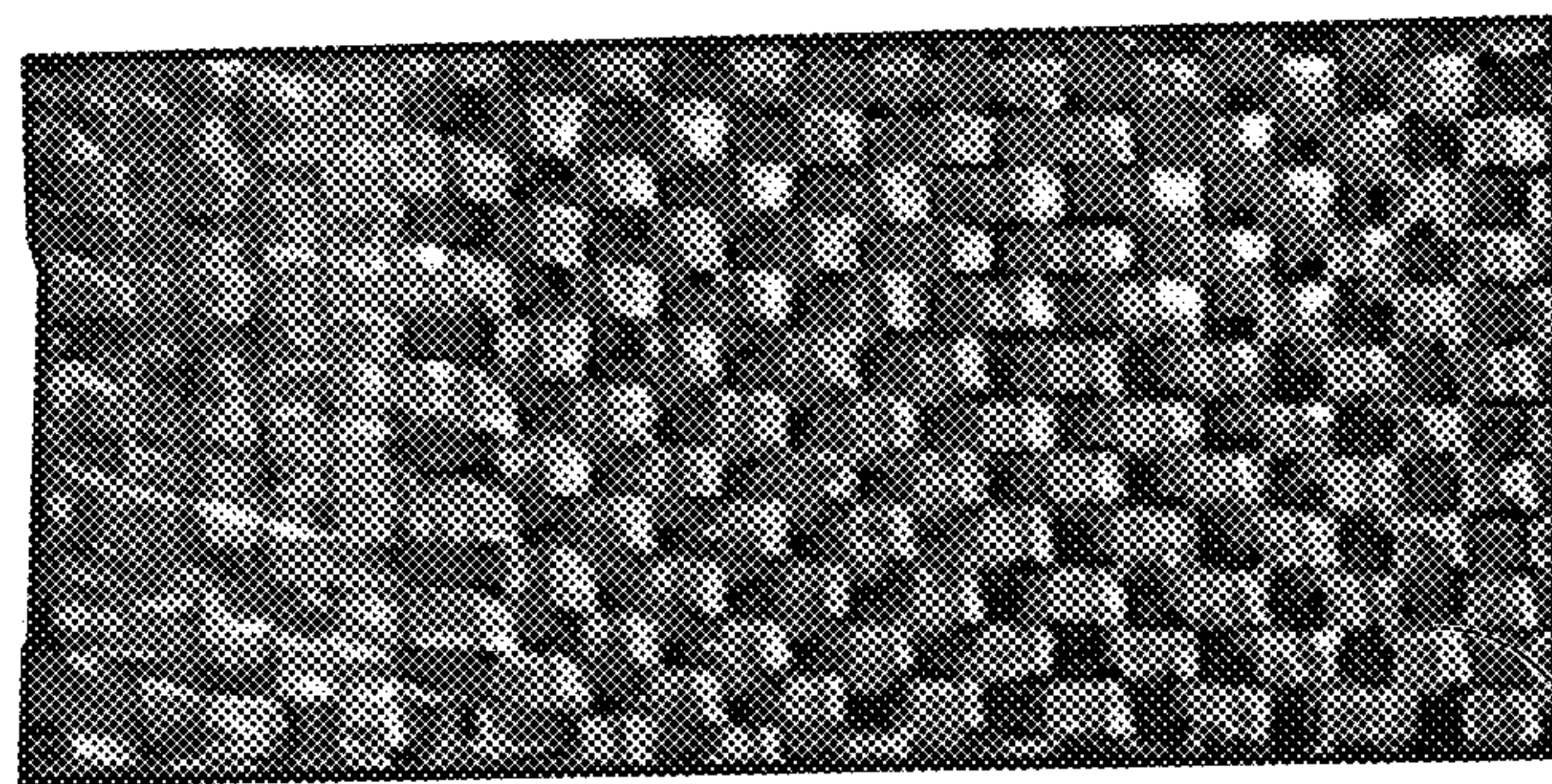


FIG. 23B

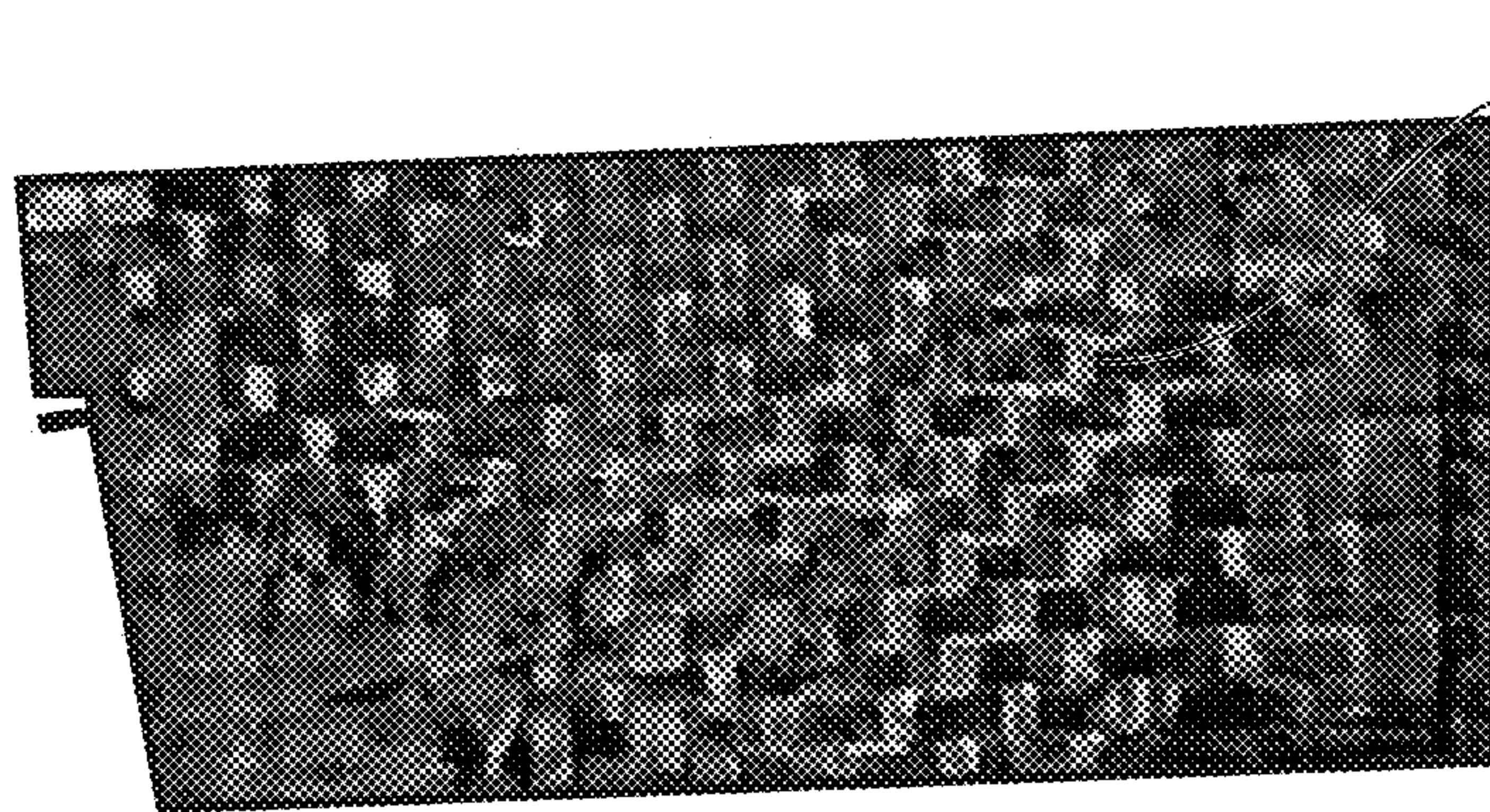


FIG. 23C

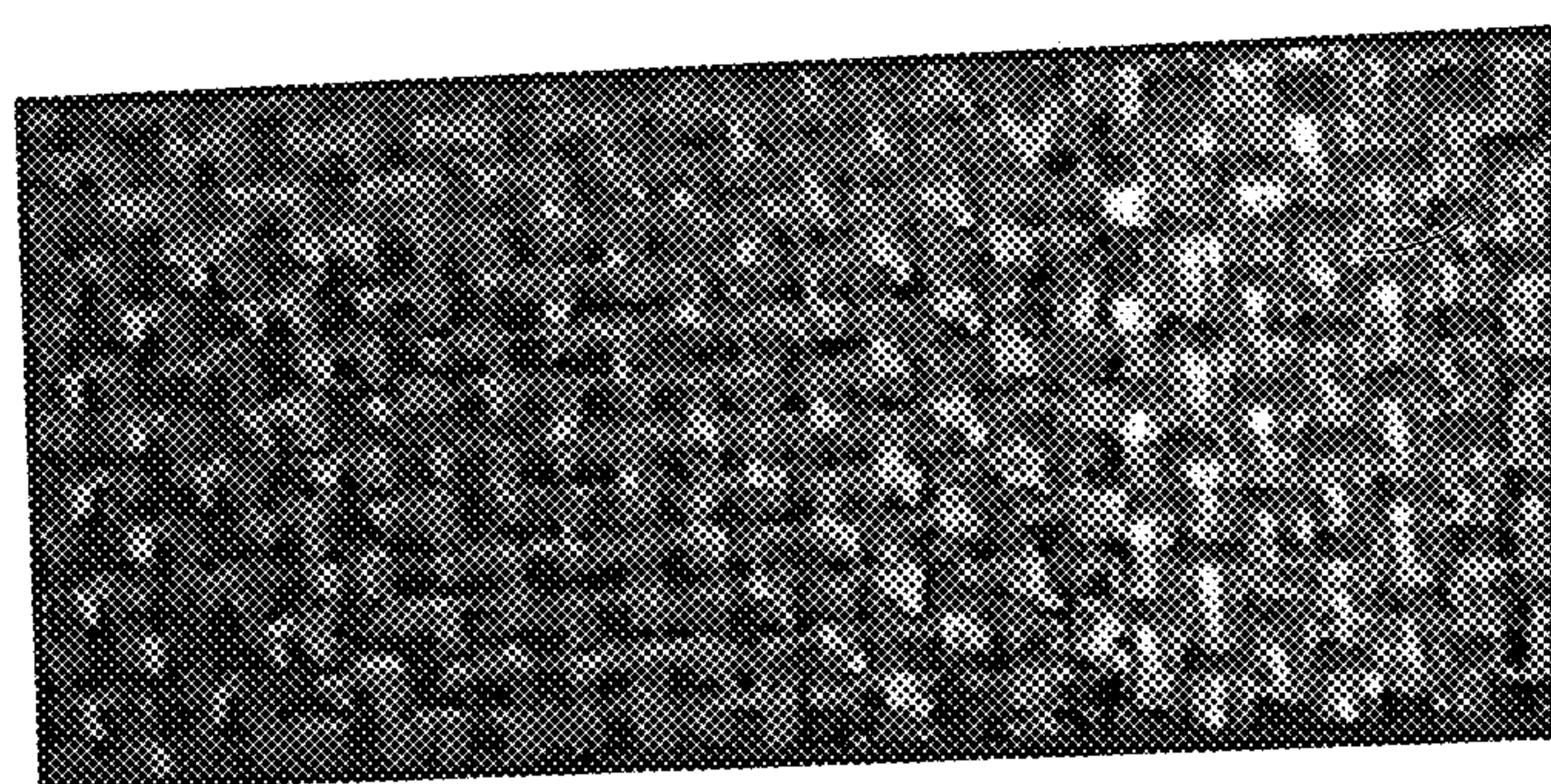


FIG. 23D



## METAL ORGANIC FRAMEWORK MODIFICATIONS OF STRUCTURAL FIBERS

### CLAIM OF PRIORITY

[0001] This patent application claims the benefit of priority, under 35 U.S.C. Section 119(e), to NI,arwan Al-Haik, U.S. Patent Application Ser. No. 63/155,579, entitled "METAL ORGANIC FRAMEWORKS MODIFICATIONS OF STRUCTURAL FIBERS," filed on Mar. 2, 2021, which is hereby incorporated by reference herein in its entirety.

### STATEMENT REGARDING FEDERALLY SPONSORED RESEARCH OR DEVELOPMENT

[0002] This invention was made with government support under award number Summer Faculty Fellowship Program (SFFP-2020) awarded by Air Force Research Laboratory (AFRL). The government has certain rights in this invention.

### BACKGROUND

[0003] Lightweight materials are often used to improve fuel efficiency of vehicles such as bicycles, cars, aircraft, etc. For example, aluminum, titanium, and fiberglass have been commonly used in vehicles to lower the vehicle's weight while maintaining a strength high enough to withstand regular operation forces. Carbon fiber has weight and strength properties that make it a good candidate for such applications. As such, carbon fiber is commonly used in vehicles of many types.

### BRIEF DESCRIPTION OF THE DRAWINGS

[0004] In the drawings, which are not necessarily drawn to scale, like numerals may describe similar components in different views. Like numerals having different letter suffixes may represent different instances of similar components. The drawings illustrate generally, by way of example, but not by way of limitation, various embodiments discussed in the present document.

[0005] FIG. 1 illustrates a schematic view of a metal-organic framework (MOF).

[0006] FIG. 2A illustrates a perspective view of spools of carbon fiber.

[0007] FIG. 2B illustrates a top view of carbon fiber wrapped around a plate.

[0008] FIG. 2C illustrates an elevation view of an assembly of plates.

[0009] FIG. 3 illustrates a perspective view of chopped carbon fibers.

[0010] FIG. 4A illustrates a scanning electron microscopy micrograph of the carbon fiber with de-sized fibers.

[0011] FIG. 4B illustrates a scanning electron microscopy micrograph of acid-treated carbon fiber.

[0012] FIG. 4C illustrates a scanning electron microscopy micrograph of MOF growth with a concentration of 0.65 M MOFs.

[0013] FIG. 4D illustrates a scanning electron microscopy micrograph of MOF growth with a concentration of 0.33 M MOFs.

[0014] FIG. 4E illustrates scanning electron microscopy micrograph of carbon nanotubes (CNTs) growth based on fibers with 0.65 M MOFs concentration.

[0015] FIG. 4F illustrates scanning electron microscopy micrograph of CNTs growth based on fibers with 0.33 M MOFs concentration.

[0016] FIG. 5A illustrates an energy-dispersive X-ray spectroscopy of carbon fibers with MOFs revealed along strong peaks of crystalline nickel.

[0017] FIG. 5B illustrates a graph of the energy-dispersive X-ray spectroscopy of FIG. 5A.

[0018] FIG. 5C illustrates an energy-dispersive X-ray spectroscopy of carbon fibers with MOFs after growing CNTs revealed along strong peaks of crystalline nickel.

[0019] FIG. 5D illustrates a graph of the energy-dispersive X-ray spectroscopy of FIG. 5C.

[0020] FIG. 6 illustrates a Raman spectrum of carbon fibers with different surface treatments.

[0021] FIG. 7A illustrates a wide-angle X-ray diffraction (WAXD) pattern of a reference sample carbon fiber.

[0022] FIG. 7B illustrates a WAXD pattern of an acid-treated carbon fiber sample.

[0023] FIG. 7C illustrates a WAX pattern of a carbon fiber sample with a MOF.

[0024] FIG. 7D illustrates a WAXD pattern of a carbon fiber sample with a reduced MOF.

[0025] FIG. 7E illustrates a WAXD pattern of a carbon fiber sample with metal-organic framework CNTs.

[0026] FIG. 8 illustrates a graph of WAXD profiles of carbon fiber samples after different surface treatments.

[0027] FIG. 9 illustrates a graph of Fourier-transform infrared (FTIR) Spectra of carbon fiber samples after different surface treatments.

[0028] FIG. 10 illustrates a zoomed-in graph of the Spectra of carbon fibers without a MOF.

[0029] FIG. 11A illustrates scanning transmission electron microscopy (STEM) micrographs of a carbon fiber sample with MOFs.

[0030] FIG. 11B illustrates STEM micrographs of a carbon fiber sample with MOFs.

[0031] FIG. 12A illustrates an Image-J® digitized of a STEM micrograph.

[0032] FIG. 12B illustrates a graph of the particle size distribution for the carbon fiber of FIG. 12A.

[0033] FIG. 13A illustrates a water meniscus around a reference carbon fiber sample.

[0034] FIG. 13B illustrates a water meniscus around an acid-treated carbon fiber sample.

[0035] FIG. 13C illustrates a water meniscus around a carbon fiber sample with MOFs.

[0036] FIG. 13D illustrates a water meniscus around a carbon fiber sample with CNTs.

[0037] FIG. 14 illustrates a graph of water contact angle vs. time for different carbon fiber samples.

[0038] FIG. 15 illustrates a graph of Thermogravimetric analysis (TGA) curves of different carbon fiber configurations.

[0039] FIG. 16A illustrates a graph of tensile tests of different carbon fiber samples having different surface treatments.

[0040] FIG. 16B illustrates a graph of average tensile strength and axial stiffness of different carbon samples with different surface treatments.

[0041] FIG. 17A illustrates micrographs of the fracture surfaces for a reference sample.

[0042] FIG. 17B illustrates micrographs of the fracture surfaces for a reference sample.

[0043] FIG. 17C illustrates micrographs of the fracture surfaces for an acid-treated sample.



[0044] FIG. 17D illustrates micrographs of the fracture surfaces for an acid-treated sample.

[0045] FIG. 18A illustrates micrographs of the fracture surfaces for a MOF fiber sample.

[0046] FIG. 18B illustrates micrographs of the fracture surfaces for a MOF carbon fiber sample.

[0047] FIG. 18C illustrates micrographs of the fracture surfaces for a carbon fiber sample with CNTs.

[0048] FIG. 18D illustrates micrographs of the fracture surfaces for a carbon fiber sample with CNTs.

[0049] FIG. 19 illustrates a graph of a dynamic mechanical analysis (DMA) measurement of the damping parameter;  $\tan(\delta)$  for different samples with different surface treatments via frequency sweep at room temperature.

[0050] FIG. 20 illustrates a graph of a DMA measurement of a temperature sweep for different samples with different surface treatments.

[0051] FIG. 21 illustrates a graph of shear lap joint stress vs. displacement for samples with different surface treatments.

[0052] FIG. 22 illustrates a graph of shear strength for samples with different surface treatments.

[0053] FIG. 23A illustrates representative images of different joint failure modes for a reference sample of carbon fiber.

[0054] FIG. 23B illustrates representative images of different joint failure modes for an acid-washed sample of carbon fiber.

[0055] FIG. 23C illustrates representative images of different joint failure modes for a MOF carbon fiber sample.

[0056] FIG. 23D illustrates representative images of different joint failure modes for a carbon fiber sample with CNTs.

#### DETAILED DESCRIPTION

[0057] Carbon fiber has weight and strength properties that make it a good candidate for use in vehicles of many types. Carbon fiber has been advanced significantly over the past 40 years. For example, many carbon fibers have been developed for their strength to weight ratios. However, carbon fiber reinforced polymeric composites (CFRPs) are prone to delamination due to insufficient interfacial properties. Several remedies have been attempted to enhance the fiber/matrix interfaces via chemical treatments or utilizing stiffer nanomaterials at the interface. However, some of these treatments are destructive in nature and others are non-scalable.

[0058] This disclosure helps to address the delamination issues by using a methodology for developing hybrid reinforcements that comprise carbon fibers and metal organic frameworks (MOFs). The growth of MOFs is both scalable, non-destructive to the carbon fibers, and tailorable to control porous morphologies of the MOF at the interface. Furthermore, this disclosure demonstrates the feasibility of utilizing the MOFs as a catalyst to grow carbon nanotubes (CNTs) on the carbon fibers. Several mechanical characterizations including tensile, dynamic mechanical analysis and shear lap joint were carried out to discern the effects of the MOF on the composite structural performance. Several improvements emanated from the MOF placement on the interface including, improving the strength, enhancing the damping parameter by 500%, increasing the glass transition temperature of the composite by 20 °C., and alleviating the shear lap joint strength by 40%.

[0059] Though the disclosure below discusses the use of carbon fiber as a substrate, glass fibers can substitute carbon fiber in any embodiment. Optionally, a mixture of glass fiber and carbon fibers can be used as a substrate.

[0060] The above discussion is intended to provide an overview of subject matter of the present patent application. It is not intended to provide an exclusive or exhaustive explanation of the invention. The description below provides further information about the present patent application.

[0061] FIG. 1 shows a schematic view of a MOF 100 illustrating an organic ligand 102 bonded to metallic nodes 104. The MOF 100 can be an organic-inorganic hybrid crystalline porous material. The MOF 100 can form an array of positively charged metal ions surrounded by organic linker molecules (e.g., organic ligand 102). The metal ions can form metallic nodes 104 that bind the arms of the organic ligand 102 together to form a repeating, cage-like structure. Due to this hollow structure, the MOF 100 can have an extraordinarily large surface area.

[0062] The MOF 100 has been previously utilized for non-structural applications. The MOF 100 with carbon materials can be used as catalysis, gas adsorption, or enhanced transport properties. For example, the MOF 100 and CNTs can form a porous catalyst used to detect urea. In another example, graphene, coated with the MOF 100, can be used as an effective water-splitting electro-catalyst. In yet another example, the MOF 100 can be a copper-based hybrid with graphene oxide. When the MOF 100 is a copper-based hybrid with graphene oxide, the MOF 100 can be used for adsorbing toxic H<sub>2</sub>S gas.

[0063] Here, the MOF 100 can modify the surface of structural carbon fibers toward amplifying the interfacial properties of CFRPs. Moreover, nickel-based MOF 100 can be used as a catalyst to grow carbon nanotubes on carbon fibers.

[0064] FIG. 2A illustrates a perspective view of a spool of carbon fiber 200. FIG. 2B illustrates a top view of the carbon fiber 200 wrapped around a plate 202. FIG. 2C illustrates an elevation view of an assembly of plates 202. FIGS. 2A-2C are discussed together below.

[0065] As shown in FIG. 2A, carbon fiber 200 can be wrapped around a core for storage and transportation. FIGS. 2A-2C show an example of carbon fiber 200 that can be bonded with the MOF 100 (FIG. 1). As shown in FIG. 2B, the carbon fiber 200 can be wrapped around the glass plate 202. Optionally, as shown in FIG. 2C multiple plates 202 can be stacked or arranged with space between the plates. One or more plates with carbon fiber wrapped around them can be immersed in a mixture to bond the MOF 100 to the carbon fiber 200.

[0066] FIG. 3 illustrates a perspective view of shredded carbon fiber 200. The carbon fiber 200 can be formed in a woven fabric, for example, and can be shredded before being submersed in the MOF 100 solutions. To prepare the MOF 100 mixture, one solution can be made by dissolving the metal ion source, nickel nitrate hexahydrate (Ni(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O) in 100 mL of methanol to yield a 0.33 M concentration. The ligands solution can be made by dissolving 2-methylimidazole (C<sub>4</sub>H<sub>6</sub>N<sub>2</sub>) in 100 mL of methanol: 0.14 M concentration. The metal ion source can use other metals, such as alkaline earth metals or transition metals, such as Cobalt Nitrate, Iron Nitrate, or the like. In other examples, the ligands solution can include any molecule that consists



of a lone pair of electrons; for example some ligand solutions can consist of an —OH, —NH<sub>3</sub>, —CO, or a combination thereof.

[0067] Various resins can be used, such as a thermoset, a thermoplastic, a phenolic, a vinyl ester, a polyimide resin, or the like. In some examples, the resin can be a polymeric resin. For example, the resin may be chosen from a diglycidyl ether of bisphenol F, a low epoxy equivalent weight diglycidyl ether of bisphenol A, a liquid epoxy novolac, a liquid aliphatic epoxy, a liquid cycloaliphatic epoxy, a 1,4-cyclohexanedimethanoldiglycidyl ether, 3,4-epoxycyclohexylmethyl-3,4-epoxy cyclohexane carboxylate, tetraglycidylmethylenedianiline, N,N,N',N'-tetraglycidyl-4,4'-methylenebisbenzenamine, a triglycidyl of para-aminophenol, N,N,N',N'-tetraglycidyl-m-xylene diamine, or a mixture thereof.

[0068] In some examples, the acrylate resin may be chosen from a methacrylate, a methyl acrylate, an ethyl acrylate, a 2-chloroethyl vinyl ether, a 2-ethylhexyl acrylate, a hydroxyethyl methacrylate, a butyl acrylate, a butyl methacrylate, or a mixture thereof.

[0069] In some examples, the methacrylate may be chosen from 2-hydroxyethyl (meth)acrylate; N-vinylpyrrolidone; N-vinylcaprolactam; acrylamide; mono- or di-N-alkyl substituted acrylamide; t-butyl acrylamide; dimethylaminoethyl acrylamide; N-octyl acrylamide; poly(alkoxyalkyl) (meth)acrylates including 2-(2-ethoxyethoxy)ethyl (meth)acrylate, 2-ethoxyethyl (meth)acrylate, 2-methoxyethoxyethyl (meth)acrylate, 2-methoxyethyl methacrylate, polyethylene glycol mono(meth)acrylates; alkyl vinyl ethers, including vinyl methyl ether; and mixtures thereof.

[0070] In some examples, the ethyl acrylate may be chosen from benzyl 2-ethyl acrylate; 2-(dimethylamino)ethyl acrylate; poly(ethyl acrylate) solution; 2-(diethylamino)ethyl acrylate; di(ethylene glycol) ethyl ether acrylate; poly(methyl methacrylate-co-ethyl acrylate); poly(tert-butyl acrylate-co-ethyl acrylate-co-methacrylic acid); 2-(4-benzoyl-3-hydroxyphenoxy)ethyl acrylate; poly(ethylene-co-ethyl acrylate); 2-[[[(butylamino)carbonyl]oxy]ethyl acrylate; 2-(9H-carbazol-9-yl)ethyl acrylate; poly(9H-carbazole-9-ethyl acrylate); poly(methyl methacrylate)-co-(9-H-carbazole-9-ethyl acrylate); ethyl 2-(hydroxymethyl)acrylate; ethyl cis-(β-cyano)acrylate; ethyl 2-(trimethylsilylmethyl)acrylate; ethyl 2-(bromomethyl)acrylate; ethyl(2e)-2-cyano-3-(1H-indolyl-3-yl)acrylate; ethyl 3-(n,n-dimethylamino)acrylate; ethyl 2-cyano-3,3-bis(methylthio)acrylate; and mixtures thereof.

[0071] In some examples, 2-ethylhexyl acrylate may be chosen from 2-Hydroxyethyl methacrylate; poly(2-hydroxyethyl methacrylate); phosphoric acid 2-hydroxyethyl methacrylate ester; and mixtures thereof.

[0072] In some examples, butyl acrylate may be chosen from tert-Butyl acrylate; poly(tert-butyl acrylate-co-ethyl acrylate-co-methacrylic acid); tert-Butyl 2-bromoacrylate, and mixtures thereof.

[0073] In some examples, butyl methacrylate may be chosen from tert-Butyl methacrylate; sec-butyl methacrylate; poly(butyl methacrylate); poly(butyl methacrylate-co-methyl methacrylate); poly(butyl methacrylate-co-isobutyl methacrylate); poly(tert-butyl methacrylate-co-glycidyl methacrylate); poly(tert-butyl methacrylate); poly(4-vinylpyridine-co-butyl methacrylate); poly[(propylmethacrylate-heptaisobutyl-PSS)-co-(t-butyl methacrylate)]; poly[(propylmethacrylate-heptaisobutyl-PSS)-co-(n-butyl methacrylate)];

diethylene glycol butyl ether methacrylate; 2-(3-(tert-butyl)-4-hydroxy-5-(5-methoxy-2H-benzo[d][1,2,3]triazol-2-yl)phenoxy)ethyl methacrylate; 3-(3-(tert-butyl)-4-hydroxy-5-(5-methoxy-2H-benzo[d][1,2,3]triazol-2-yl)phenoxy)propyl methacrylate; 2-(Dimethylamino)ethyl methacrylate; and mixtures thereof.

[0074] In some examples, the polysiloxane resin may be chosen from poly[dimethyl siloxane-co-methyl(3-hydroxypropyl)siloxane]-graft-poly(ethylene glycol) methyl ether, poly[dimethyl siloxane-co-methyl(stearoyloxyalkyl)siloxane], poly(methyl hydrogen siloxane), 1,1,3,3-tetramethyldisiloxane, and mixtures thereof.

[0075] FIG. 4A illustrates a scanning electron microscopy micrograph of the carbon fiber **200** with de-sized fibers. The carbon fibers **200** utilized can be de-sized, plain woven, structural, polyacrylonitrile (PAN) grade, Thornel-650, with 3k bundles. The fibers can be de-sized inside a quartz tube furnace, at 550° C. for 30 mins under a nitrogen-induced inert environment. To allow the formation of active sites (—COOH), the fibers can be cut into 6"×6" squares and can be treated in a diluted mix (2:1) of de-ionized (DI) water/HNO<sub>3</sub> acid for 24 hours. The fibers can be washed thoroughly with DI water till a pH of 7.0 was achieved, followed by drying in an oven at 100° C. for 24 hours. A reference sample of **390** can be made of the carbon fiber **200**.

[0076] The peel-off, as shown in FIG. 4A, can indicate that de-sizing the T650 fibers under an inert thermal environment was a success. The carbon fiber **200** can be treated with the diluted HNO<sub>3</sub> solution dissolved in these de-sized peels; nitric acid is well known for introducing the carboxylic group (—COOH) to the fibers' surface. The acid treatment can also play an etching effect revealing several fibrils at the surface of the carbon fiber **200**.

[0077] FIG. 4B illustrates a scanning electron microscopy micrograph of acid-treated carbon fiber sample **400**. As shown in FIG. 4B, the surface of acid-treated carbon fiber **400** can be smooth and does not contain the peel-off demonstrated on the surface of the carbon fiber sample **400** in FIG. 4A.

[0078] FIG. 4C illustrates scanning electron microscopy micrograph of a MOF coated carbon fiber sample **402** having a concentration of 0.65 M MOF **100** (FIG. 1). The MOF **100** porous structure shown in FIG. 4C can be formed utilizing a highly concentrated solution of the metal source. As shown in FIG. 4C pores **404** can be separated by thin sheet walls **406**. The thin sheet walls **406** can be, for example, MOF **100**. The thin sheet walls **406** can help highlight the pattern of the MOF **100** formed. Several factors can play roles in creating the different patterns of the MOF **100**, such as the metal source concentration, the ligand utilized (e.g., organic ligand **102**), the growth time, and the nature of the substrate. For example, decreasing the concentration of the MOF **100** of a metal-organic framework carbon fiber **402** can make the thin sheet walls **406** disappear.

[0079] FIG. 4D illustrates scanning electron microscopy micrograph of another reduced MOF coated carbon fiber sample **408** a concentration of 0.33 M MOF **100**. As shown in FIG. 4D, decreasing the concentration of the MOF **100** can make the thin sheet walls **406** (shown in FIGS. 4C) disappear.

[0080] FIG. 4E illustrates scanning electron microscopy micrograph of a carbon fiber sample with CNTs **410** growth based on fibers with 0.65 M MOF **100** concentration. As



shown in FIG. 4E, the dense MOF **100** (shown in FIG. 4C), can lead to dense CNTs growth that can enlarge fibers of the carbon fiber **200** to have diameters from 6,8  $\mu\text{m}$  to almost 30.0  $\mu\text{m}$ . The dense growth of the carbon fiber **200** on the carbon fiber sample with CNTs **410** can hinder the ability of the resin to wet the fibers and penetrate through the bundles.

[0081] FIG. 4F illustrates a scanning electron microscopy micrograph of a carbon fiber sample with CNTs **412** based on fibers with 0.33 M MOF **100** concentration. The less-dense MOF **100** (shown in FIG. 4C) can be used to make less-dense CNT growth resulting in more uniform and less-dense growth of the CNT. The uniform and less-dense development can allow better flowability of epoxy between the fibers of the carbon fiber sample with CNTs **412**.

[0082] FIG. 5A illustrates energy-dispersive X-ray spectroscopy of a sample **500** that can be made from the carbon fiber **200** with the MOF **100**. (e.g., MOF coated carbon fiber sample **402** from FIG. 4C or the reduced MOF coated carbon fiber sample **408** from FIG. 4D. FIG. 5B illustrates a graph of the energy-dispersive X-ray spectroscopy of FIG. 5A. FIG. 5C illustrates energy-dispersive X-ray spectroscopy of a sample **502** that can be made from the carbon fiber **200** with the MOF **100** after growing CNTs (e.g., the carbon fiber sample with CNTs **410** of FIG. 4E the carbon fiber sample with CNTs **412** from FIG. 4F. FIG. 5D illustrates a graph of the energy-dispersive X-ray spectroscopy of FIG. 5C. FIGS 5A-5D will be discussed together below.

[0083] As shown in FIGS. 5A and SB, the energy-dispersive X-ray spectroscopy of the sample **500** can reveal strong peaks of crystalline nickel. As shown in FIGS. 5C and 5D, the carbon peak can become stronger because of the added CNTs atoms to the sample **502** and the MOF **100**. Moreover, the Ni spots can become more contrasted because the reduction process, before CNTs grown of the sample **502**, can assist in removing oxides that might have been developed during the MOF **100** depositions.

[0084] FIG. 6 illustrates a Raman spectra of carbon fibers with different surface treatments. As shown in FIG. 6, the reference sample of the **390** can reveal a strong disordered graphitic peak (D;  $1360\text{ cm}^{-1}$ ) (D peak), and a much weaker ordered graphitic peak (G;  $1580\text{ cm}^{-1}$ ) (G peak). The lesser referred to peak can be at roughly  $1500\text{ cm}^{-1}$ , which can indicate amorphous carbon. The acid-treated carbon fiber sample **400** can show reduced peaks of intensities of both D and G peaks. Further, the G peak can be broadened and shifted to a higher frequency because of the  $-\text{COOH}$  functional groups on the surface. The reduction in the G band intensity can indicate a reduced number of graphitic layers. As also shown in FIG. 6 with reference to the sample **500**, the deposition of the MOF **100** on the carbon fiber **200** can mask the D and G peaks due to higher coverage of metal crystals. The spectra can also measure the absorption of the MOF **100**. As also shown in FIG. 6 with reference to the sample **502**, upon growing CNTs on a surface of the carbon fiber **200**, a strong D-band at  $1350\text{ cm}^{-1}$  and G-band at  $1595\text{ cm}^{-1}$  can be apparent. An intensity ratio,  $I_D/I_G$  can be more noticeable for the sample **502**, indicating a high degree of crystallinity due to the presence of the CNTs.

[0085] FIG. 7A illustrates a WAXD pattern of the reference sample of carbon fiber **390**. FIG. 7B illustrates a WAXD pattern of an acid-treated carbon fiber sample **400**, FIG. 7C illustrates a WAXD pattern of the MOF coated carbon fiber sample **402**. FIG. 7D illustrates a WAXD pattern of the reduced MOF carbon fiber sample **408**. FIG.

7E illustrates a WAXD pattern of the carbon fiber sample with CNTs **412**. FIGS. 7A-7E will be discussed concurrently below.

[0086] As shown in FIGS. 7A-7E show that the patterns can indicate a diffuse anisotropic scattering. FIGS. 7A and 7B show that the the reference sample of carbon fiber **390** and the acid-treated carbon fiber sample **400** show a fan-like scattering typical for PAN-based carbon fibers. The needle-like pores on the fibers of the examples of FIGS. 7A and 7B can cause the scattering pattern. As shown in FIGS. 7A and 7B, the reference sample of carbon fiber **390** and acid-treated carbon fiber sample **400** can exhibit identical patterns, which indicates that the acid treatment did not alter the crystallinity of the fibers of the carbon fiber **200**. The intensity distribution in the reciprocal space includes disks centered at the origin of that space with disks normal parallel to the principal axis of the pores. Disc thickness can be inversely related to the length of the pores. The fan-like shape can include a superposition of disks with finite thicknesses.

[0087] FIG. 8 illustrates a graph of WAXD profiles of carbon fiber fibers after different surface treatments. As shown in FIG. 8, PAN fibers can display a turbostratic carbon peak. The turbostratic carbon peak beyond  $2\theta=25^\circ$  can be typical for turbostratic fibers and it appears in all samples as the fibers serve as a substrate for the different coatings.

[0088] The samples (e.g., **402**. or **408**) can have a different scattering pattern with many crystalline peaks can emerge. The different scattering pattern can possess a large unit cell because its scattering peaks do not match those of the Ni (111) crystallographic plane of NiO (200) and (111) planes. When reduced at  $550^\circ\text{C}$ . under inert environment, the samples (**402** or **408**) can lose several of these peaks and shows a noticeable peak at  $2\theta=39^\circ$ . The peak of  $2\theta=39^\circ$  of the MOF **100** can be indicative of the NiO (111) plane. The sample **502** can display a maximum peak at  $26.2^\circ$  that can correspond to a wall-to-wall periodicity in multiple wall CNTs. The maximum peak of the samples (10 or 12) can be a characteristic of the (002) reflection from the graphitic layer.

[0089] The WAXS patterns and graphs shown in FIG. 7E and 8 confirm that the sample **502** can be multiwalled CNT's because a single wall CAT comprises a single graphitic layer. In contrast, to a single wall CNT, the diffraction shown in FIGS. 7E and 8 of the sample **502** can indicate that multiple planes of graphitic layers are present. Interspacing between the graphitic shells can be computed from:

$$d = \frac{2\pi}{q_{\max}} = 3.4\text{Å} \quad [1]$$

where the scattering vector is given by:

$$q = 4\pi\sin\left(\frac{2\theta}{\lambda}\right) \quad [2]$$

where  $\lambda$  is wavelength.

[0090] FIG. 9 illustrates a graph of FUR Spectra of the carbon fiber **200** after different surface treatments. FIG. 10



illustrates a zoomed-in graph of the Spectra of the carbon fiber **200** without MOF **100**. FIGS. **9** and **10** will be discussed together below.

[0091] As shown in FIGS. **9** and **10**, the reference sample of carbon fiber **390** does not show discernible or typical peaks due to the lack of any functional groups. The acid-treated carbon fiber sample **400** can show new peaks. For example, the acid-treated carbon fiber sample **400** can have groups C—O—C, C—O—N, or C—N that can appear between  $140(\text{ )cm}^{-1}$  and  $1000\text{ cm}^{-1}$ . As shown in FIGS. **9** and **10**, these peaks for the acid-treated carbon fiber sample **400** can be stronger than peaks of the reference sample of carbon fiber **390**. These stronger peaks for the acid-treated carbon fiber sample **400** can indicate that oxygen functional groups were introduced through oxidation by the nitric acid. The peaks between  $3200$  and  $3500\text{ cm}^{-1}$  can be associated with OH group stretching vibrations, which can indicate water presence due to insufficient drying of the carbon fiber **200**.

[0092] FIGS. **9** and **10** can also indicate that the MOF **100** were successfully deposited onto the carbon fiber **200** for the sample **500** because the sample **500** can have many peaks that are not present in other surface treatments of the carbon fiber **200**. A first peak around  $3400\text{ cm}^{-1}$  can be caused by the O—H group coupling to Ni(ii). A second group of peaks at around  $1590\text{ cm}^{-1}$  and  $1390\text{ cm}^{-1}$  can be associated with a stretching of coordinated carboxylate ( $\text{—COO}^-$ ) and symmetric stretching mode of coordinated carboxylate, respectively. A third group of peaks, between  $7500\text{—}800\text{ cm}^{-1}$  can ascribe to the O—Ni—O vibrations while thus around  $1050$  and  $850\text{ cm}^{-1}$  can confirm the C—N and C—H bonds, respectively.

[0093] FIG. **11A** illustrates STEM micrographs of the MOF coated carbon fiber sample **402**. FIG. **11B** illustrates STEM micrographs of the reduced MOF carbon fiber sample **408** after reduction under Nitrogen. FIGS. **11A** and **11B** will be discussed together below.

[0094] The pre-reduction morphology of the MOF coated carbon fiber sample **402**, shown in FIG. **11A**, shows that the carbon fiber **200** can have interwoven flakes of the MOF **100**. The interwoven flakes of the MOF **100** can result in the thin sheet walls **406** (FIG. **4C**). The reduced-MOF carbon fiber sample **408**, shown in FIG. **11B**, shows that the reduced MOF carbon fiber sample **408** can have particulate shapes with broad particle size distribution. The larger particles morphology of the carbon fiber **200**, as shown in FIG. **11B**, can suggest that the reduced MOF **100** of the carbon fiber **200** can suggest that the carbon fiber **200** can comprise agglomerates of much smaller particles.

[0095] STEM-EDS mode can be used to find the composition of the MOF coated carbon fiber sample **402** and the reduced-MOF carbon fiber sample **408**.

TABLE 1

Element	MOF post reduction (%)				Element	MOF pre reduction (%)			
	C	N	O	Ni		C	N	O	Ni
Average	73.08	2.64	4.35	19.93	Average	80.26	3.35	5.05	11.33
St. Deviation	4.54	0.33	0.38	5.14	St. Deviation	5.68	1.43	2.05	2.64

[0096] As shown in Table 1, reducing the MOF-coated carbon fiber sample **402** to the reduced-MOF carbon fiber sample **408** can decrease the carbon content. The reduction

in carbon content in the reduced-MOF carbon fiber sample **408** can be caused by partial decomposition of the MOF **100** (FIG. **1**) in the reduced-MOF carbon fiber sample **408**. As also shown in Table 1, reducing the MOF coated carbon fiber sample **402** to the reduced-MOF carbon fiber sample **408** can decrease the oxygen content and increase the nickel content. The reduction of oxygen and increase in nickel can suggest the removal of oxides in making the reduced-MOF carbon fiber sample **408**. Thus, because the removal of oxides is preferable to aid in the catalytic growth of CNTs, the reduced-MOF carbon fiber sample **408** can be better suited for catalytic growth of CNTs.

[0097] FIGS. **12A** and **12B** will be discussed below concurrently. FIG. **12A** illustrates an Image-J® digitized STEM micrograph of the carbon fiber sample with CNTs **412**. FIG. **12B** illustrates a graph of the particle size distribution for the carbon fiber **200** of the carbon fiber sample with CNTs **412** of FIG. **12A**. As shown in FIGS. **12A** and **12B**, the carbon fiber **200** of the carbon fiber sample with CNTs **412** can have no dominant size of the reduced particles. The particles of the carbon fiber **200** of the carbon fiber sample with CNTs **412** can range from a few nm to almost 250 nm. Such a wide range of particle sizes can lead to different CNTs, or even carbon nanofibers, diameters.

[0098] FIG. **13A** illustrates a water meniscus around the reference sample of carbon fiber **390**. FIG. **13B** illustrates a water meniscus around the acid-treated carbon fiber sample **400**. FIG. **13C** illustrates a water meniscus around the reduced MOF coated carbon fiber sample **408**. FIG. **13D** illustrates a water meniscus around the carbon fiber sample with CNTs **412**, FIGS. **13A-13D** will be discussed together below.

[0099] FIGS. **13A-13D** compare a water meniscus on a surface of the carbon fiber **200** with different surface treatments. Each of FIGS. **13A-13D**, can be representative of a sample **100** seconds after a sample contacts water. The meniscus values shown in FIG. **13A-13D** can be considered qualitatively as the sample geometry was changing with time due to capillary induced densification of the samples, which can be visible specially in the reference samples. Loose fibers in the samples aggregates can be due to capillary effect during the immersion in water. However, as shown in FIG. **13C**, the reduced MOF coated carbon fiber sample **408** can retain almost a constant diameter throughout the immersion process while the other three configurations can exhibit necking in the diameter. As shown in FIG. **13D**, the carbon fiber sample with CNTs **412** can have the largest diameter due to an added thickness of the CNTs coating and due to nanoscale gaps between the CNTs, which can minimize the water capillary effect.

[0100] FIG. **14** illustrates a graph of water contact angle vs. time for different samples of the carbon fiber **200**. Each

of the samples shown in FIG. **14** can show quasi-instantaneous meniscus forms evident by a drop of a contact angle from  $90^\circ$  to anywhere between  $65^\circ$  and  $45^\circ$ . The reduced



MOF coated carbon fiber sample **408** can have the greatest drop. The bend of each curve in FIG. **14** can show an amount of time for water to completely fill each sample. The reference sample of carbon fiber **390** can have the shortest amount of time to completely fill with water, and the acid-treated carbon fiber sample **400** can have the second shortest amount of time to completely fill with water. Each of the reduced MOF coated carbon fiber sample **408** and the carbon fiber sample with CNTs **412** can take much longer to fill with water. The delay in filling with water in the reduced MOF coated carbon fiber sample **408** and the carbon fiber sample with CNTs **412** can be because the water needs to force the air out of the porous structures of the samples, This multiscale capillary effect of the reduced MOF coated carbon fiber sample **408** and the carbon fiber sample with CNTs **412** can explain the increased time needed to fill the samples with water. After the water replaces the air within a sample, the contact angle can attain its static value. The reference sample of the carbon fiber **390** can have a contact angle of  $46.31^\circ$ . The acid-treated carbon fiber sample **400** can have a contact angle of  $35.07^\circ$ . The reduced MOF coated carbon fiber sample **408** can have a contact angle of  $33.43^\circ$ . The carbon fiber sample with CNTs **412** can have a contact angle of  $34.52^\circ$ .

[0101] FIG. **15** illustrates a graph of TGA curves of different configurations of the carbon fiber **200**. TGA curves can show a thermal stability of different samples of the carbon fiber **200** with different surface treatments. As shown in FIG. **15**, the reference sample of the carbon fiber **390** can attain a 10-percent weight loss of the original weight in a nitrogen environment. In the reference sample of the carbon fiber **390**, the lack of a sizing agent or other surface coatings can lead to fast and significant degradation.

[0102] The reduced MOF coated carbon fiber sample **408** can display the fastest degradation rate up to  $225^\circ\text{C}$ ., which can be likely due to insufficient drying as the samples were soaked separately in water during acid-treatment and in methanol during the growth of the MOF **100**. Thus, the reduced MOF coated carbon fiber sample **408** was likely exposed to more moisture than any of the other samples. For the reduced MOF coated carbon fiber sample **408**, the degradation of the sample past  $22.5^\circ\text{C}$ . can slow down as compared to the acid-treated carbon fiber sample **400** and the reference sample of carbon fiber **390**. At about  $650^\circ\text{C}$ . the degradation of the reduced MOF coated carbon fiber sample **408** can increase. The increase in the degradation of the reduced MOF coated carbon fiber sample **408** can suggest that the sample is decomposing.

[0103] The carbon fiber sample with CNTs **412** can have the lowest degradation of all of the samples The carbon fiber sample with CNTs **412** can have a good graphitization structure which can result in less temperature degradation.

[0104] FIG. **16A** illustrates a graph of tensile tests of different carbon fiber samples with different surface treatments. FIG. **16B** illustrates a graph of average tensile strength and axial stiffness of different carbon samples with different surface treatments. FIGS. **16A** and **16B** will be discussed together below.

[0105] As shown in FIGS. **16A** and **16B**, all of the samples can exhibit linear elastic behavior and the stiffness for all of the samples remained almost unaltered by the surface treatments. The stiffness remained the same for all of the samples

because stiffness can be a volumetric property that can depend on the core of the fiber, rather than the surface of the fiber.

[0106] The carbon fiber sample with CNTs **412** can have a slight improvement on the stiffness as compared to other samples. The improvements of stiffness of the carbon fiber sample with CNTs **412** can likely be due to annealing at elevated temperatures which can sinter the MOF **100** under the CNTs, which can lead to a stiffer phase.

[0107] The acid-treated carbon fiber sample **400** can have a slight increase of strength (2.1%) when compared to the reference sample of carbon fiber **390**. The small increase in strength of the acid-treated carbon fiber sample **400** can likely be attributed to the relatively low acid concentration acid (30%) used to prepare the samples. The low acid concentration acid was used to help control density of MOF **100** growth in samples. In some examples, the acid etching of the acid-treated carbon fiber sample **400** can cause damage to fibers of the carbon fiber **200**, which can result in a decrease in strength of a sample. [OWN] The MOF coated carbon fiber sample **402** can increase the strength by 11% as compared to the reference sample of carbon fiber **390**. The MOF **100** on the surface of the carbon fiber **200** on the MOF coated carbon fiber sample **402** can increase the surface area and provide anchoring between the fiber and the matrix. Defects induced by acid-etching can be filled with the MOF **100**, as the  $-\text{COON}$  grouped induced by the acid etching and the ligand (e.g., organic ligand **102**) in the MOF **100** can play a role in anchoring the NI-MOF structure to the activated sites on the surface of the carbon fiber **200**. The MOF **100** can assist in better load transition from the matrix to the fiber, and better load distribution on the fiber surface. The better load transition from matrix to fiber and better load distribution on the fiber surface can increase the load-bearing capacity of the sample, which can result in enhanced strength.

[0108] The carbon fiber sample with CNTs **412** can improve the strength by 5.5% when compared to the reference sample of carbon fiber **390**. The elevated temperature needed to grow the CNTs can induce some damage to the carbon fiber **200**. This damage to the carbon fiber **200** can be in addition to the damage caused by the acid-etching of the carbon fiber **200**. Growing fibers on the surface of the carbon fiber **200** can lead to modest strength improvements, The dense blanket of CNTs can hinder the flow of the epoxy from reaching the underlying carbon fiber, which can be limit the quality of adhesion. Having the MOE **100** on the surface of the carbon fiber **200** beneath the CNTs can help improve the flow of epoxy and can help improve the strength of a sample.

[0109] FIG. **17A**. illustrates micrographs of the fracture surfaces for a reference sample of carbon fiber **390**. FIG. **17B** illustrates micrographs of the fracture surfaces for a reference sample of carbon fiber **390**. FIGS. **17A** and **17B** will be discussed together.

[0110] As seen in FIGS. **17A** and **17B**, the fracture of the composite can be dictated by fiber breakage. The fiber breakage was likely caused by the absence of a strong interface between the fiber and the matrix in the reference sample of carbon fiber **390**. This fiber breakage can be shown by the clean fiber surface in FIG. **17B**. The fiber breakage can lead to an unstoppable crack propagating across the carbon fiber **200**.

[0111] FIG. **17C** illustrates micrographs of the fracture surfaces for the acid-treated carbon fiber sample **400**. FIG.



17D illustrates micrographs of the fracture surfaces for the acid-treated carbon fiber sample 400. FIGS. 17C and 17D will be discussed together.

[0112] As shown in FIGS. 17C and 17D, activating the surface of the carbon fiber 200 in the acid-treated carbon fiber sample 400 did not alter the pattern of fiber failure. However, as shown in FIGS. 17C and 17D, the surface of the carbon fiber 200 in the acid-treated carbon fiber sample 400 can improve the adhesion between layers of the carbon fiber 200, which can be shown by the residue of the epoxy on the surface of the carbon fiber 200.

[0113] FIG. 18A illustrates micrographs of the fracture surfaces for the reduced MOF coated carbon fiber sample 408. FIG. 18B illustrates micrographs of the fracture surfaces for the reduced MOF coated carbon fiber sample 408. FIGS. 18A and 18B will be discussed together.

[0114] As shown in FIGS. 18A and 18B, the MOF 100 on the reduced MOF coated carbon fiber sample 408 can act as crack deterrents or retardants. Cracks that initiate in the matrix can get deflected along the axis of the carbon fiber 200 on a wavy trajectory as shown in FIG. 18A. The MOF 100 (FIG. 1) of the MOF coated carbon fiber sample 402 can stay intact through the failure as shown in FIG. 18B.

[0115] FIG. 18C illustrates micrographs of the fracture surfaces for the carbon fiber sample with CNTs 412. FIG. 18D illustrates micrographs of the fracture surfaces for the carbon fiber sample with CNTs 412. FIGS. 18C and 18D will be discussed together.

[0116] As shown in FIGS. 18C and 18D, the CNTs on the surface of the carbon fiber 200 of the carbon fiber sample with CNTs 412 can hinder crack propagation along the fiber axis and the matrix in between. As shown, adjacent fibers can sometimes not undergo longitudinal cracks as well. The rough surfaces shown in FIG. 18D, can indicate good adhesion between the fibers and the matrix.

[0117] For both the reduced MOF coated carbon fiber sample 408 and the carbon fiber 200 of the carbon fiber sample with CNTs 412 the samples can have little or no cracks between the fibers. The patterned structure of the MOF 100 (FIG. 1) and the formation of CNTs can promote mechanical interlocking between the nano reinforcement species and the epoxy matrix. The mechanical interlocking between the nano reinforcement species and the epoxy matrix can constitute strong impediments to crack propagation and can increase the strength of the samples.

[0118] FIG. 19 illustrates a graph of a DMA measurement of the damping parameter;  $\tan(\delta)$  for different samples with different surface treatments via frequency sweep at room temperature.

[0119] Interfacial bonding, stick-slip friction, and nano-fillers are different mechanisms that can affect the damping capacity of carbon-fiber-reinforced composites. As shown in FIG. 19, the reference sample of carbon fiber 390 can have the lowest damping parameter. The acid-treated carbon fiber sample 400 can show as much as 200% at 87 Hz improvement as compared to the reference sample of carbon fiber 390. The acid-treated carbon fiber sample 400 likely has this improvement because the acid-etching can enhance the surface area of the carbon fiber 200, which can allow for larger contact with the epoxy.

[0120] The addition of MOT 100 to the carbon fiber 200 in the MOF coated carbon fiber sample 402 can increase the surface area even more, and the weak epoxy-metal interfacial bonding can allow for more energy dissipation. Thus,

for the MOF coated carbon fiber sample 402, the damping parameter can be increased as high as 500% as compared to the reference sample of carbon fiber 390.

[0121] The growth of CNTs on the carbon fiber sample with CNTs 412 can increase the surface area tremendously because of the high aspect ratio of the CNTs. This increased surface area can result in significant friction between the CNTs and the epoxy, which can lead to significant energy dissipation. The carbon fiber sample with CNTs 412 can attain an improvement of the damping coefficient by up to 425% as compared to the reference sample of carbon fiber 390.

[0122] FIG. 20 illustrates a graph of a DMA measurement of a temperature sweep for different samples with different surface treatments.

[0123] DMA measurements can also elucidate how different surface treatments of the carbon fiber 200 can affect the glass transition temperature  $T_g$ , of the carbon fiber reinforced composites. The glass transition temperature,  $T_g$ , can be measured at the peak of  $\tan(\delta)$  curve. As shown in FIG. 20, the reference sample of carbon fiber 390 can have a glass transition temperature,  $T_g$ , of about 120°. The acid-treated carbon fiber sample 400 can show no significant change of a glass transition temperature,  $T_g$ , when compared to the reference sample of carbon fiber 390. The reduced MOF coated carbon fiber sample 408 can have an elevated glass transition temperature, as compared to the reference sample of carbon fiber 390, of 140° C. The carbon fiber sample with CNTs 412 can have an even higher glass transition temperature,  $T_g$ , of 150° C. Because the reduced MOF coated carbon fiber sample 408 and the carbon fiber sample with CNTs 412 have an increased glass transition temperature,  $T_g$ , the interface embodied by a new phase between the fiber and matrix can play a role in a shift of the glass transition temperature,  $T_g$ . During a glass transition, epoxy molecular chains can absorb thermal energy and start moving. However, having the MOF 100 or CNTs on the surface of the carbon fiber 200 (e.g., the reduced MOF coated carbon fiber sample 408 and the carbon fiber sample with CNTs 412) can slow down the movement relating to the glass transition. Thus, more energy (e.g., temperature) can be necessary to overcome these nanoscale hurdles caused by the MOF 100 and CNTs on the carbon fiber 200. The measurement of glass transition temperature,  $T_g$ , can offer a venue to, qualitatively, prove the effects of fiber surface treatment on the mechanical performance of epoxy-based composites.

[0124] FIG. 21 illustrates a graph of shear lap joint stress vs. displacement for samples with different surface treatments. FIG. 22 illustrates a graph of shear strength for samples with different surface treatments. FIGS. 21 and 22 will be discussed below.

[0125] The reference sample of carbon fiber 390 can have the lowest joint shear strength. The acid-treated carbon fiber sample 400 can have a shear strength that is 10% higher than that of the reference sample of carbon fiber 390. The enhanced surface roughness of the acid-treated carbon fiber sample 400 can increase the shear strength of the acid-treated carbon fiber sample 400.

[0126] The MOF coated carbon fiber sample 402 can have a joint shear strength that is 40% higher than the reference sample of carbon fiber 390. The nanosheet structures of the MOF coated carbon fiber sample 402 can furnish an extended contact area and a mechanical interlocking of



epoxy between the pores of the MOF coated carbon fiber sample **402** can be responsible for the increased joint shear strength.

[0127] The carbon fiber sample with CNTs **412** can improve the shear strength by 4.4% as compared to the reference sample of carbon fiber **390**. The minute increase in strength can indicate that the epoxy did not wet the CNTs, which is likely due to the density of the CNTs in the carbon fiber sample with CNTs **412**.

[0128] FIGS. **23A-23D** will be discussed together. FIG. **23A** illustrates representative images of different joint failure modes observed for the reference sample of carbon fiber **390**. FIG. **23B** illustrates representative images of different joint failure modes observed for the acid-washed sample of carbon fiber **400**. FIG. **23C** illustrates representative images of different joint failure modes observed for the MOF carbon fiber sample **408**. FIG. **23D** illustrates representative images of different joint failure modes observed for the carbon fiber sample with CNTs **412**.

[0129] As shown in FIG. **23A**, the reference sample of carbon fiber **390** can have a stock-break failure (e.g., the sample broke outside the joint area). As shown in FIG. **23B**, the acid-treated carbon fiber sample **400** can experience a cohesive failure. The cohesive failure can be evidenced by the adhesive residue left behind on both parts of the acid-treated carbon fiber sample **400** one of which is shown in FIG. **23B**). The surface functionality of the acid-treated carbon fiber sample **400** can be responsible for improvements in shear strength.

[0130] As shown in FIG. **23C**, the MOF coated carbon fiber sample **408** can have a fiber-tear failure. The fiber-tear failure can be indicative of the reinforced fiber on both ruptured surfaces. The tearing of the fiber can be indicative of significant shear bearing capability. As shown in FIG. **23D**, the carbon fiber sample with CNTs **412** can show that the epoxy did not wet the CNTs well, which can be due to the density of the CNTs. Thus, there can be a thin-layer cohesive failure, leaving some adhesive dusting on one of the joint surface, and a thicker adhesive dusting on the other joint surface.

#### NOTES AND EXAMPLES

[0131] The following, non-limiting examples, detail certain aspects of the present subject matter to solve the challenges and provide the benefits discussed herein, among others.

[0132] Example 1 is a reinforced carbon composite comprising: a carbon substrate; and a metal organic framework bonded to the carbon substrate.

[0133] in Example 2, the subject matter of Example 1 optionally includes wherein the carbon substrate comprises carbon nanolubes,

[0134] In Example 3, the subject matter of any one or more of Examples 1-2 optionally include wherein the carbon substrate comprises carbon fibers.

[0135] In Example 4, the subject matter of Example 3 optionally includes wherein the carbon substrate comprises woven polyacrylonitrile carbon fibers.

[0136] In Example 5, the subject matter of any one or more of Examples 1-4 optionally include wherein the metal organic framework is a reaction product of a metal ion source and a ligands solution.

[0137] In Example 6, the subject matter of Example 5 optionally includes wherein the ion source comprises nickel nitrate hexahydrate,

[0138] In Example 7, the subject matter of any one or more of Examples optionally include -methylimidazole.

[0139] Example 8 is a carbon fiber reinforced polymeric composite comprising: a first layer comprising: a carbon substrate; and a metal organic framework bonded to the carbon substrate; a second layer comprising: a carbon substrate; and a metal organic framework bonded to the carbon substrate; and a resin adhered to the first layer and the second layer.

[0140] Example 9, the subject matter of Example 8 optionally includes wherein the resin comprises a polymeric resin.

[0141] In Example 10, the subject matter of any one or more of Examples optionally include wherein the resin comprises one or more resins chosen from a. thermoset resin, a thermoplastic resin, a phenolic resin, a vinyl ester resin, a polyimide resin, an epoxy resin, an acrylic resin, an acrylate resin, a cyanoacrylate resin, cyano-urethane resin, a polysiloxane resin, or a. mixture thereof.

[0142] In Example 11, the subject matter of any one or more of Examples 8-10 optionally include wherein the carbon substrate is made of carbon nanotubes.

[0143] In Example 12, the subject matter of any one or more of Examples 8-11 optionally include wherein the carbon substrate comprises carbon fibers.

[0144] In Example 13, the subject matter of Example 12 optionally includes wherein the carbon substrate comprises woven polyacrylonitrile carbon fibers.

[0145] In Example 14, the subject matter of any one or more of Examples 8-13 optionally include wherein the metal organic framework is a reaction product of a metal ion source and a ligands solution.

[0146] In Example 15, the subject matter of Example 14 optionally includes wherein the ion source comprises nickel nitrate hexahydrate.

[0147] In Example 16, the subject matter of Example 15 optionally includes wherein the ligand solution comprises 2-methylimidazole.

[0148] Example 17 is a method of forming a carbon fiber reinforced polymeric composite, the method comprising: forming a first substrate and a second substrate by: forming a metal ion solution by dissolving a. metal ion source in methanol; forming a ligands solution by dissolving 2-methylimidazole in methanol; mixing the metal ion solution and the ligands solution; immersing the carbon fiber in the mixture; and washing and drying the carbon fiber; and adhering the first substrate to the second substrate using a resin.

[0149] In Example 18, the subject matter of Example 17 optionally includes wherein the ligand solution comprises 2-methylimidazole.

[0150] In Example 19, the subject matter of any one or more of Examples 17-18 optionally include wherein de-sizing includes heating the carbon fibers to **550** degrees Celsius for 30 minutes.

[0151] In Example 20, the subject matter of any one or more of Examples 17-19 optionally include wherein the ion source is nickel nitrate hexahydrate.

[0152] in Example 21, the subject matter of any one or more of Examples 17-20 optionally include de-sizing carbon fibers; soaking the carbon fibers in a water and acid bath; washing the carbon fibers; and drying the carbon fibers.



**[0153]** In Example 22, the subject matter of any one or more of Examples 17-21 optionally include wherein soaking the carbon fibers in a water and acid bath includes soaking the carbon fibers in a mixture of de-ionized water and HNO<sub>3</sub> acid

**[0154]** In Example 23, the subject matter of any one or more of Examples 17-22 optionally include stirring the metal ion solution, magnetically; and stirring the ligands solution, magnetically.

**[0155]** Example 24 is a method of forming carbon nanotubes using a metal organic framework, the method comprising: producing a metal organic framework by: forming a metal ion solution by dissolving a metal ion source in methanol; forming a ligands solution by dissolving 2-methylimidazole in methanol; and mixing the metal ion solution and the ligands solution; and forming carbon nanotubes within or around the metal organic framework.

**[0156]** In Example 25, the subject matter of Example 24 optionally includes-rnethylimidazole.

**[0157]** In Example 26, the subject matter of any one or more of Examples 24-25 optionally include wherein the ion source is nickel nitrate hexahydrate.

**[0158]** In Example 27, the subject matter of any one or more of Examples 24-26 optionally include stirring the metal ion solution, magnetically; and stirring the ligands solution, magnetically.

**[0159]** In Example 28, the subject matter of any one or more of Examples 24-27 optionally include wherein forming carbon nanotubes within the metal organic framework includes: inserting the metal organic framework into a gas reactor; removing excessive oxides from the metal organic framework by flowing a gas mixture of H<sub>2</sub>/N<sub>2</sub> over the metal organic framework at 550 degrees Celsius for two hours; flushing the gas reactor with N<sub>2</sub> gas; filling the reactor with C<sub>2</sub>H<sub>4</sub>/H<sub>2</sub>/N<sub>2</sub> gas; and heating the reactor to 550 degrees Celsius for 30 minutes

**[0160]** Example 29 is a reinforced glass composite comprising: a fiberglass substrate; and a metal organic framework bonded to the fiberglass substrate.

**[0161]** Example 30 is a reinforced composite comprising: a substrate; and a metal organic framework bonded to the substrate.

**[0162]** In Example 31, the subject matter of Example 30 optionally includes wherein the substrate is fiberglass or carbon fiber.

**[0163]** Example 32 is a three-dimensional printed component formed of carbon fiber reinforced polymeric composite as shown and described herein.

**[0164]** Example 33 is a method for fabricating a carbon fiber reinforced polymeric composite as shown and described herein.

**[0165]** Example 34 is a method of forming carbon fiber nanotubes supported by a metal organic framework as shown and described herein.

**[0166]** Example 35 is a fiber reinforced polymeric composite comprising: a substrate; and a metal organic framework bonded to the substrate.

**[0167]** In Example 36, the apparatuses or method of any one or any combination of Examples 1-35 can optionally be configured such that all elements or options recited are available to use or select from.

**[0168]** The above detailed description includes references to the accompanying drawings, which form a part of the detailed description. The drawings show, by way of illus-

tration, specific embodiments in which the invention can be practiced. These embodiments are also referred to herein as “examples.” Such examples can include elements in addition to those shown or described. However, the present inventors also contemplate examples in which only those elements shown or described are provided. Moreover, the present inventors also contemplate examples using any combination or permutation of those elements shown or described. (or one or more aspects thereof), either with respect to a particular example (pr one or more aspects thereof), or with respect to other examples (or one or more aspects thereof) shown or described herein.

**[0169]** In the event of inconsistent usages between this document and any documents so incorporated by reference, the usage in this document controls. In this document, the terms “including” and “in which” are used as the plain-English equivalents of the respective terms “comprising” and “wherein.” Also, in the following claims, the terms “including” and “comprising” are open-ended, that is, a system, device, article, composition, formulation, or process that includes elements in addition to those listed after such a term in a claim are still deemed to fall within the scope of that claim.

**[0170]** The above description is intended to be illustrative, and not restrictive. For example, the above-described examples (or one or more aspects thereof) may be used in combination with each other. Other embodiments can be used, such as by one of ordinary skill in the art upon reviewing the above description. The Abstract is provided to comply with 37 C.F.R. § 1.72(b), to allow the reader to quickly ascertain the nature of the technical disclosure. It is submitted with the understanding that it will not be used to interpret or limit the scope or meaning of the claims. Also, in the above Detailed Description, various features may be grouped together to streamline the disclosure. This should not be interpreted as intending that an unclaimed disclosed feature is essential to any claim. Rather, inventive subject matter may lie in less than all features of a particular disclosed embodiment. Thus, the following claims are hereby incorporated into the Detailed Description as examples or embodiments, with each claim standing on its own as a separate embodiment, and it is contemplated that such embodiments can be combined with each other in various combinations or permutations. The scope of the invention should be determined with reference to the appended claims, along with the full scope of equivalents to which such claims are entitled.

1. A reinforced carbon composite comprising: a carbon substrate; and a metal organic framework bonded to the carbon substrate.
2. The reinforced carbon composite of claim 1, wherein the carbon substrate comprises carbon nanotubes.
3. The reinforced carbon composite of claim 1, wherein the carbon substrate comprises carbon fibers.
4. The reinforced carbon composite of claim 3, wherein the carbon substrate comprises woven polyacrylonitrile carbon fibers.
5. The reinforced carbon composite of claim 1, wherein the metal organic framework is a reaction product of a metal ion source and a ligands solution,
6. The reinforced carbon composite of claim 5, wherein the ion source comprises nickel nitrate hexahydrate.



7. The reinforced carbon composite of claim 5, wherein the ligand solution comprises 2-methylimidazole.

8. A carbon fiber reinforced polymeric composite comprising:

a first layer comprising:

a carbon substrate; and

a metal organic framework bonded to the carbon substrate;

a second layer comprising:

a carbon substrate; and

a metal organic framework bonded to the carbon substrate; and

a resin adhered to the first layer and the second layer.

9. The carbon fiber reinforced polymeric composite of claim 8, wherein the resin comprises a polymeric resin.

10. The carbon fiber reinforced polymeric composite of claim 8, wherein the resin comprises one or more resins chosen from a thermoset resin, a thermoplastic resin, a phenolic resin, a vinyl ester resin, a polyimide resin, an epoxy resin, an acrylic resin, an acrylate resin, a cyanoacrylate resin, cyano-urethane resin, a polysiloxane resin, or a mixture thereof.

11. The carbon fiber reinforced polymeric composite of claim 8, wherein the carbon substrate is made of carbon nanotubes.

12. The carbon fiber reinforced polymeric composite of claim 8, wherein the carbon substrate comprises carbon fibers.

13. The carbon fiber reinforced polymeric composite of claim 12, wherein the carbon substrate comprises woven polyacrylonitrile carbon fibers.

14. The carbon fiber reinforced polymeric composite of claim 8, wherein the metal organic framework is a reaction product of a metal ion source and a ligands solution.

15. The carbon fiber reinforced polymeric composite of claim 14, wherein the ion source comprises nickel nitrate hexahydrate.

16. The carbon fiber reinforced polymeric composite of claim 15, wherein the ligand solution comprises 2-methylimidazole.

17. A method of forming a carbon fiber reinforced polymeric composite, the method comprising:

forming a first substrate and a second substrate by:

forming a metal ion solution by dissolving a metal ion source in methanol;

forming a ligands solution by dissolving 2-methylimidazole in methanol;

mixing the metal ion solution and the ligands solution;

immersing the carbon fiber in the mixture; and

washing and drying the carbon fiber; and

adhering the first substrate to the second substrate using a resin.

18. The method of claim 17, wherein the ligand solution comprises 2-methylimidazole.

19. The method of claim 17, wherein de-sizing includes heating the carbon fibers to 550 degrees Celsius for 30 minutes.

20. The method of claim 17, wherein the ion source is nickel nitrate hexahydrate.

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