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*CeO₂-CePO₄ and Ag@CeO₂-CePO₄ nanocomposites from *Penaeus semisulcatus* for heavy metals sensing, UV shielding and cytotoxic applications*

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CeO₂-CePO₄ and Ag@CeO₂-CePO₄ nanocomposites from *Penaeus semisulcatus* for heavy metals sensing, UV shielding and cytotoxic applications

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ABSTRACT

CeO₂-CePO₄ and Ag@CeO₂-CePO₄ nanocomposites (NCs) were prepared using *Penaeus semisulcatus* and characterized by various analytical techniques. The NCs exhibited spherical morphology, with average particle sizes measuring 2.36 nm, 3.30 nm and 29.64 nm for the samples air dried, annealed at 550 °C and 800 °C respectively, and the Selected Area Electron Diffraction patterns from HR-TEM revealed their crystalline nature. In 3d X-ray Photoelectron Spectra (XPS) of NCs, the highest binding energies were observed at 916.46 eV (3d_{3/2}) and 898.16 eV (3d_{5/2}) with spin-orbit splitting of 18.3 eV. The X-Ray Diffraction (XRD) patterns of the NCs showed peaks typical of both CeO₂ and CePO₄ phases. The Fourier Transform-Infra Red (FT-IR) spectrum contained bands due to Ce-O stretching at 560 cm⁻¹ and PO₄³⁻ groups at 1008 and 711 cm⁻¹, which was further substantiated by Raman spectrum, which had an intense peak at 468 cm⁻¹ corresponding to F_{2g} band of CeO₂ and another peak at 976 cm⁻¹ due to PO₄³⁻ tetrahedron. In the Ultraviolet-Visible Diffuse Reflectance Spectra (UV-Vis DRS), the broad absorption peak at 352 nm and minor peaks around 300 nm confirmed the formation of CeO₂-CePO₄ NCs. The green synthesized NCs were excellent fluorescence sensors for Pb²⁺, Cd²⁺ and Hg²⁺ in a wide range of concentrations, from 5.0x10⁻⁹ M to 4.0x10⁻⁵ M. The concentration Vs. emission fitting plot had three distinctive linear ranges. At a low concentration range (5x10⁻⁹ to 3.5x10⁻⁷ M), the limit of detection for Pb²⁺, Cd²⁺ and Hg²⁺ was found to be 2.41x10⁻⁸, 2.56x10⁻⁸ and 3.11x10⁻⁸ M. The sensitive detection of a mixture of heavy metal ions in ppb levels by biocompatible NCs gained insight in treating aqueous industrial wastes. Further, NCs showed cytotoxicity and UV shielding properties.

1. Introduction

Nanoparticles (NPs) have brought about a revolution in technology by finding applications in diverse fields including medicine, military, aerospace, cosmetics, textile, electronics, catalysis, energy generation and storage. Metal nanoparticles find applications as nanomedicines, solar cells, supercapacitors, electroluminescent displays, catalysts, biosensors and electrochemical sensors (Singh et al., 2021). Metal oxide

nanoparticles find utility in regenerative antioxidation, catalytic converters, antimicrobial coatings, ultraviolet filters in sunscreen, fluorescent probes, light emitting diodes and *in vivo* cell imaging (Nadeem et al., 2020).

Nanoparticles are typically produced through physical, chemical and biological approaches. Biosynthesis is referred to as 'green synthesis' as it is environmentally benign, less toxic and cost-effective. Biosynthetic routes generally utilize microorganisms and plant components as

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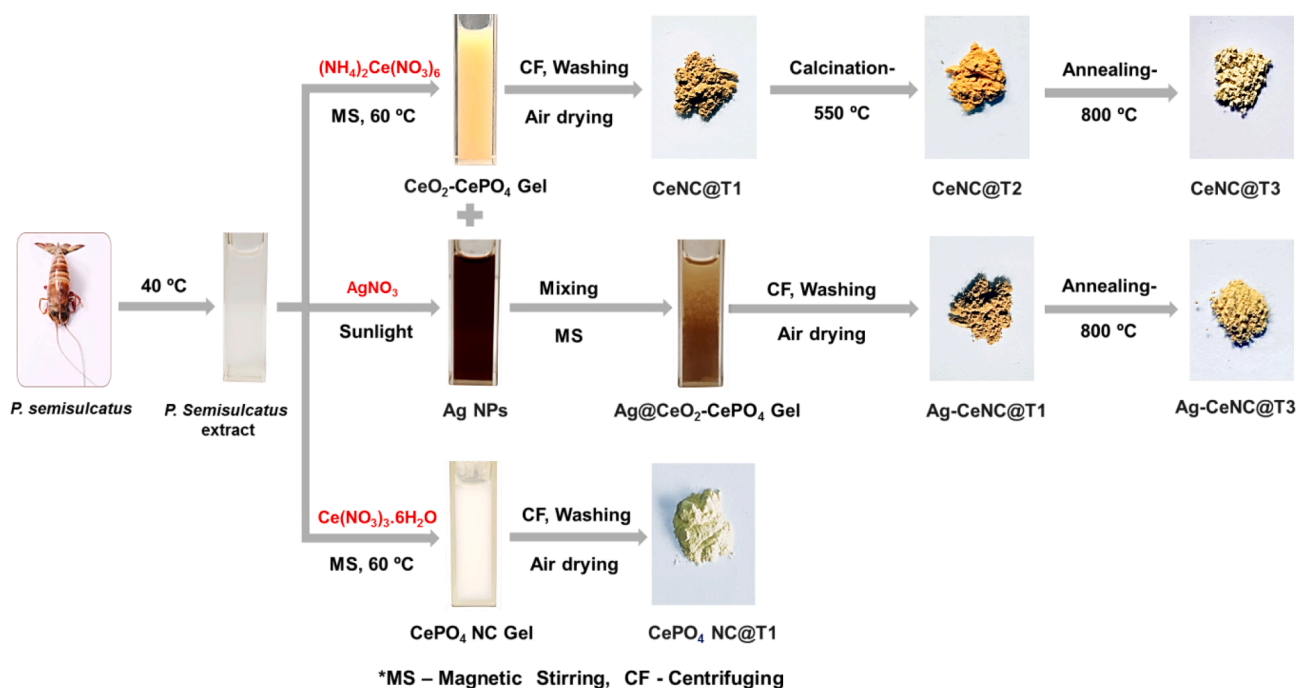
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Scheme 1. Schematic representation of *P. semisulcatus* extract mediated synthesis of NCs.

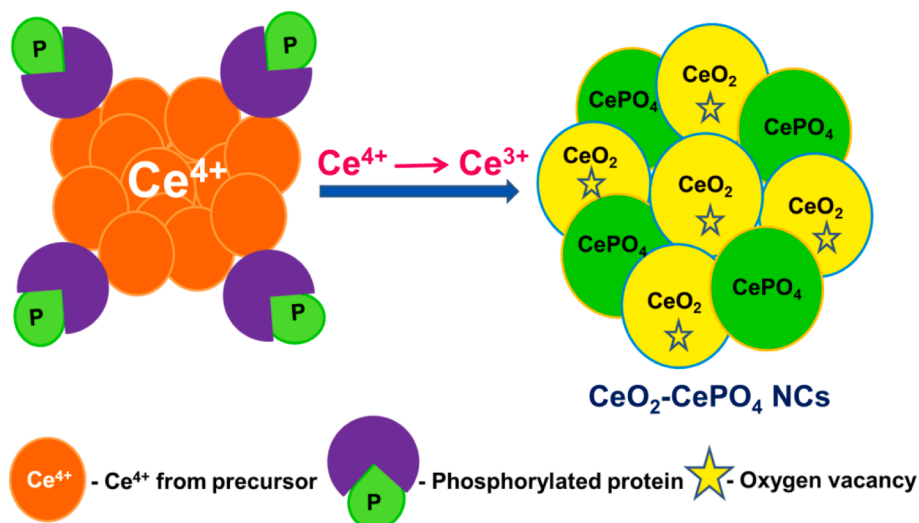


Fig. 1. Possible mechanism for the one stretch synthesis of $\text{CeO}_2\text{-CePO}_4$ NCs.

reducing (Devi et al., 2016) and capping agents (Priya and Peter, 2021), and they also serve as templates. Prokaryotic bacteria, yeast and fungi are some microorganisms that have been used in biosynthesis. Further, plant bio metabolites from plant leaves (Devi et al., 2022), roots, stem (Priya et al., 2016), seeds, etc. have been used for the syntheses of metal oxides, bimetallic and noble metals NPs (Subramanian et al., 2021). The unique optical, mechanical, surface, thermal, chemical, electrical, biological and physical characteristics of nanomaterials, synthesized by green methods make them superior to those synthesized by chemical methods (Fauzi et al., 2022).

Cerium oxide nanoparticles (CeO_2 NPs), with fluorite crystalline structure, stand out as unique among the diverse array of metal oxide nanoparticles. Affordability, chemical stability, $\text{Ce}^{3+}/\text{Ce}^{4+}$ redox stability, biocompatibility, oxygen storage potential and non-hazardous nature are some fascinating properties of ceria nanoparticles which enhance their suitability for diverse applications (Sun et al., 2020). They

find manifold applications in agriculture, fuel oxidation catalysis, supported catalysis (Sheeba et al.), corrosion protection, enzyme mimicking, neuroprotection, radiotherapy, bioremediation, automotive exhaust treatment, semiconductors, hydrogen storage materials, pharmacological agents, bio scaffolding, drug delivery and sensing. CeO_2 NPs have low surface area, a restricted range of visible light absorption and a propensity for high electron-hole recombination and are generally overcome by doping and formation of hetero junction (Singh et al., 2020).

Cerium phosphate nanoparticles (CePO_4 NPs), which exhibit high UV absorptivity and low photocatalytic activity, have been proven to be efficient UV filters. They find applications in cosmetics, nuclear waste storage materials, automobile catalytic poisons, laser technology, fluorescence labels and as proton-conducting membranes in fuel cells (Peter et al., 2022). CePO_4 nanotubes are efficient biosensors and have bio-electrochemical applications. CePO_4 NPs have been generally

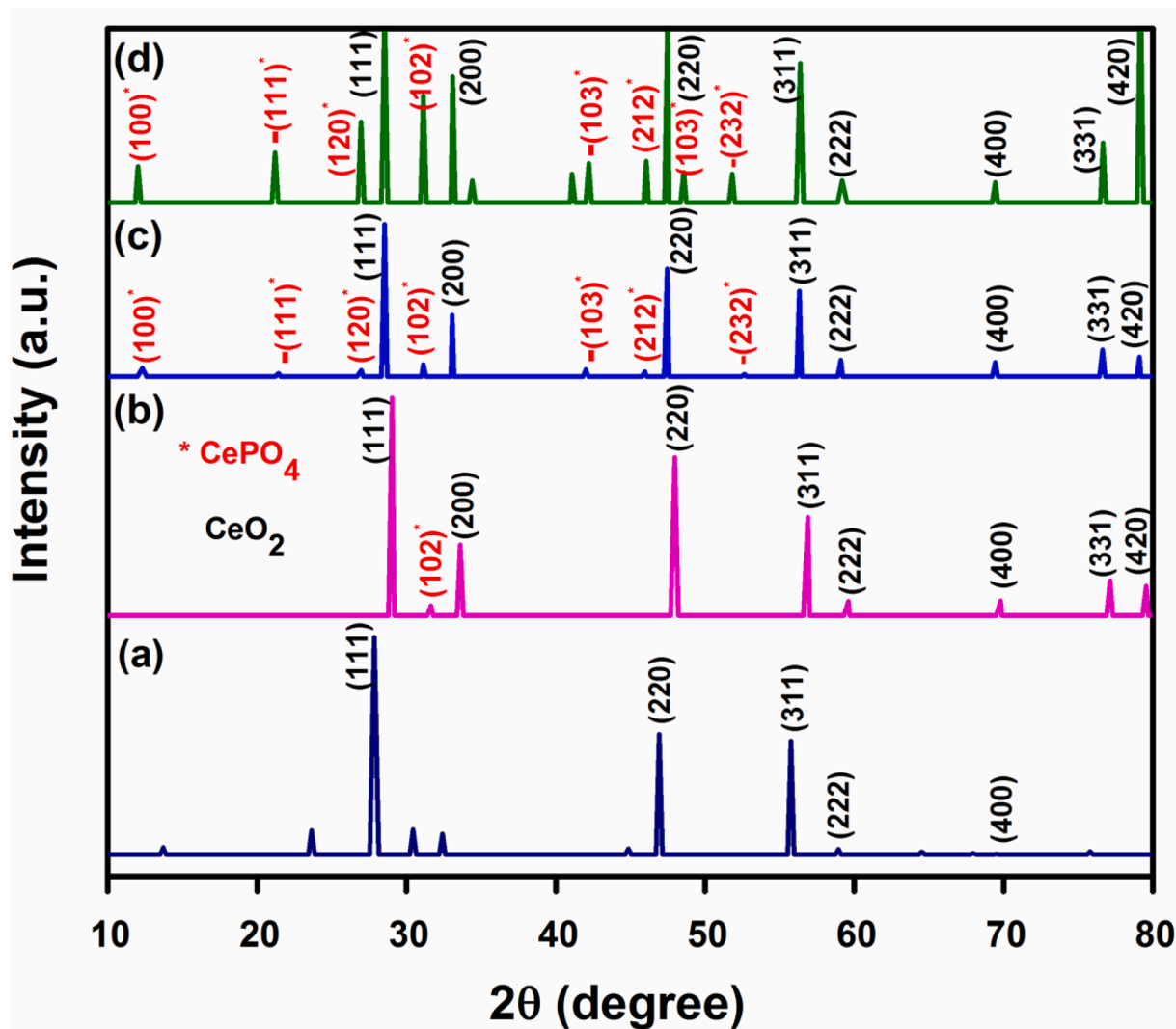


Fig. 2. XRD patterns of (a) CeNC@T1, (b) CeNC@T2, (c) CeNC@T3 and (d) Ag-CeNC@T3.

synthesized through hydrothermal method, sol-gel route, coprecipitation, solvothermal, sonochemical and reverse microemulsion techniques (De and Serra, 2013). It exists in two crystal forms, viz., monoclinic and hexagonal. The monoclinic form formed at higher temperatures has a longer luminescent lifetime. Transformation of CePO_4 from hexagonal to monoclinic phase can be induced by adjusting the initial $\text{PO}_4^{3-}/\text{Ce}^{3+}$ ratio. The luminescent properties of CePO_4 can be improved by doping with metal ions and decorating with CeO_2 nanomaterials (Vinothkumar et al., 2018). The method of preparation and calcination temperature influence their optical properties. The photoluminescence of CePO_4 NPs is associated with surface defects and oxygen vacancies, which create energy levels bridging between the Ce 4f state and the O 2p valence band. These defects serve as emission centers and do not exhibit any quenching effect. Optical studies of CePO_4 NPs reveal that they have direct band gap, allowing an electron to directly radiate a photon (Bouddouch et al., 2021).

CeO_2 - CePO_4 nanocomposites (NCs) have been produced by hydrothermal, coprecipitation and chemical routes (Fijolek and Wolski, 2022) and they are remarkably pronounced for their reversible switching between Ce^{3+} and Ce^{4+} oxidation states in nano dimension. The reduction of Ce^{4+} to Ce^{3+} leads oxygen deficiencies (positively charged) within the CeO_2 lattice. The electron transfer between CePO_4 and CeO_2 enhances their luminescence properties (Noor et al., 2021). Though biosynthesis of metal oxide and metal NPs has been well explored using plant

resources, very little attention has been drawn towards the utility of sustainable and edible marine resources (Yosri et al., 2021). A thorough study of literature reveals that Ag and Au NPs have been biosynthesised from marine algae (Fawcett et al., 2017), seaweeds (Devi et al., 2022), sea cucumber (*Holothuria scabra*) (Sepahvand et al., 2020), jellyfish (*Nemopilema nomurai*), marine invertebrate (*polychaete*), seahorse (*Hippocampus* spp.), Se NPs from marine oyster (*Abalone viscera*) (Yosri et al., 2021); ZnO NPs from *Penaeus semisulcatus* (Ishwarya et al., 2019) and CeO_2 NPs from marine oyster (Safat et al., 2021). There is only one report on the biosynthesis of CeO_2 - CePO_4 NCs using leaf extract of *Artocarpus heterophyllus* (Noor et al., 2021).

Penaeus semisulcatus is abundant in bioactive constituents, encompassing essential proteins like haemocyanin, lectin, crustin, as well as a spectrum of amino acids, fatty acids, vitamins, and minerals (Preetham et al., 2020). Studies have indicated a notable presence of amino acids such as arginine, histidine, threonine, lysine, methionine, isoleucine, leucine, phenylalanine, and valine (El-Gendy et al., 2018). Additionally, the composition comprises saturated fatty acids like stearic and palmitic acid, the monounsaturated oleic acid and polyunsaturated fatty acids such as docosahexaenoic acid, cholesterol, eicosapentaenoic acid, omega-3 fatty acids, and total omega-6 fatty acids. Vitamin content encompasses A, B1, B2, B3, B6, B12, C, E and K. Furthermore, the extract contains minerals such as iron, phosphorus, potassium, magnesium, sodium, calcium, zinc, and copper (AlFaris et al., 2022).

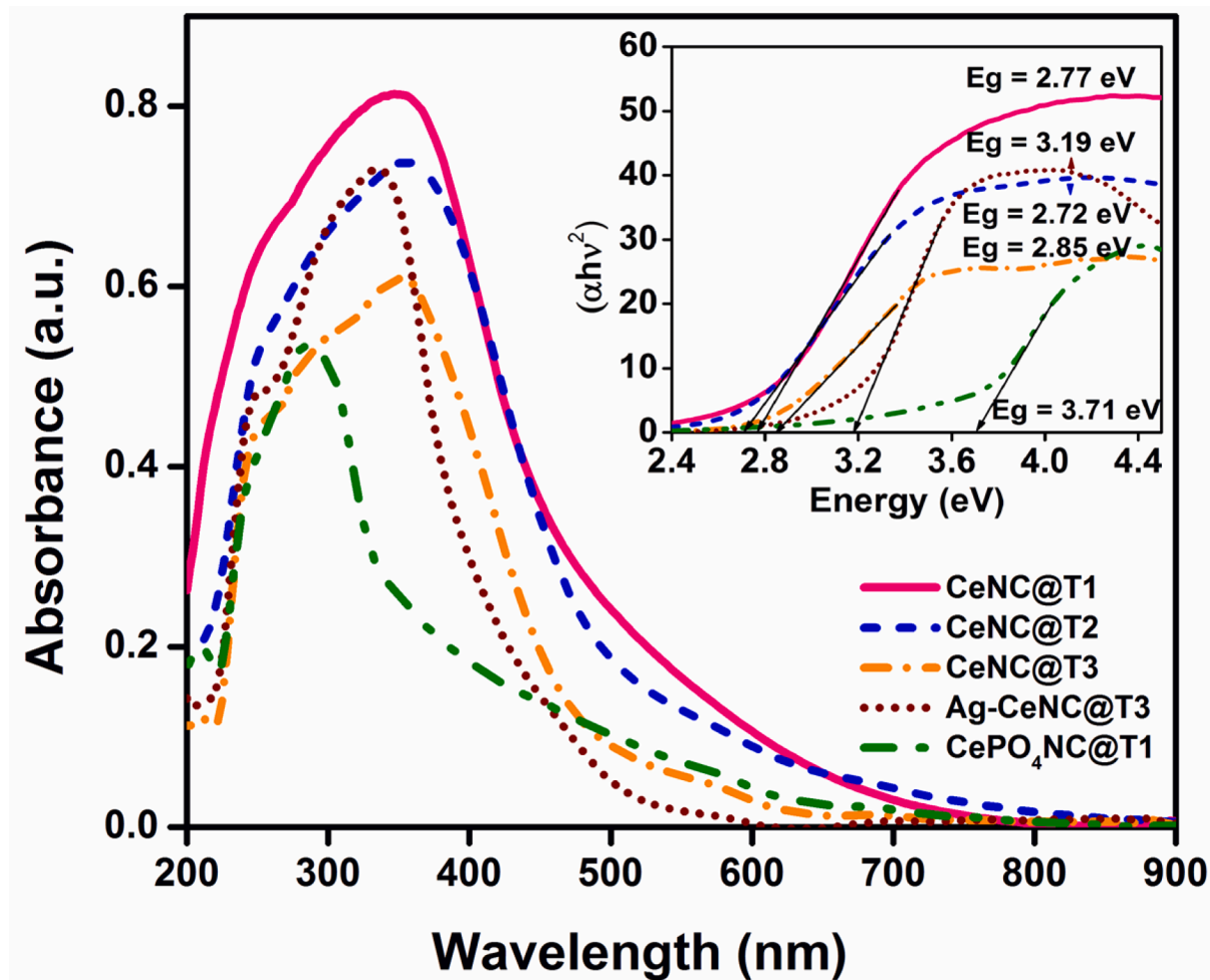


Fig. 3. UV-Vis DRS spectra of NCs. Inset: Band gap energy calculation using Tauc plot.

After conducting an in-depth review of existing literature, it has been found that edible marine shrimp have not previously been utilized in the fabrication of CeO₂-CePO₄ NCs. This study marks a pioneering effort in presenting an environmentally friendly, swift and cost-effective single-step biosynthetic approach for the creation of CeO₂-CePO₄ NCs and silver doped CeO₂-CePO₄ NCs (Ag@CeO₂-CePO₄ NCs) utilizing an aqueous extract of *P. semisulcatus*, without externally introducing phosphate precursors. This methodology sets the stage for a novel approach in the biosynthesis of various types of bio-nanocomposites.

This study has three primary goals: (i) harnessing the potential of the edible marine organism *P. semisulcatus* for the bio-fabrication of CeO₂-CePO₄ NCs, (ii) utilizing these NCs for fluorescence sensing of heavy metals and (iii) investigating the UV shielding and cytotoxic properties of the NCs for potential applications in sunscreens.

2. Experimental section

2.1. Materials

(NH₄)₂Ce(NO₃)₆ and (Ce(NO₃)₃·6H₂O) (99.9 %, AR Grade) were purchased from Sisco Research Laboratories. AgNO₃ (≥99.5 %, AR Grade) was procured from Sigma-Aldrich. EDTA disodium salt, C₁₀H₁₄N₂Na₂O₈·2H₂O (AR grade, 99.0 %), CoCl₂·6H₂O, CuSO₄·6H₂O, NiCl₂·6H₂O (SRL, 99.0 %), CrCl₃ anhydrous (99.9 %), FeCl₃ anhydrous (99.9 %), HgCl₂ (99.5 %), Cd(NO₃)₂·4H₂O (98.0 %), Pb(NO₃)₂ (99.0 %) and (NH₄)₂Fe(SO₄)₂·6H₂O (99.9 %) were acquired from Sigma Aldrich. The chemicals were utilized in their original form without any

modifications.

2.2. Collection of *P. semisulcatus* and preparation of extract

Fresh and live *P. semisulcatus* were collected by trawl fishing from the Gulf of Mannar in Thoothukudi, Tamil Nadu, India. Shrimp shells and heads were removed and cleaned with copious amounts of water. Finally, the flesh part of the *P. semisulcatus* was dissected into tiny pieces and a 30 g portion of the flesh was weighed and then gently warmed to 40 °C in a water bath with 100 mL of bidistilled water for a duration of 10 min, followed by gradual cooling to ambient temperature. The resulting extract, obtained by filtering the supernatant through Whatman No. 41 filter paper, was utilized in the production of NCs.

2.3. Biosynthesis of CeO₂-CePO₄ NCs

To a solution of ceric ammonium nitrate (2.7412 g/100 mL, 0.05 M), an equal volume of the *P. semisulcatus* extract was slowly added dropwise (1:1 v/v) under magnetic stirring, followed by heating at 60 °C for 1 h. The yellow gel formed was further stirred in the magnetic stirrer for another 2 h at 1500 rpm. The protein-rich CeO₂-CePO₄ NCs was then centrifuged, extensively washed with water and air dried (labelled as CeNC@T1). Powdered CeNC@T1 was annealed at 550 °C for 3 h (labelled as CeNC@T2) and again annealed at 800 °C for 2 h (labelled as CeNC@T3) to obtain CeO₂-CePO₄ NCs.

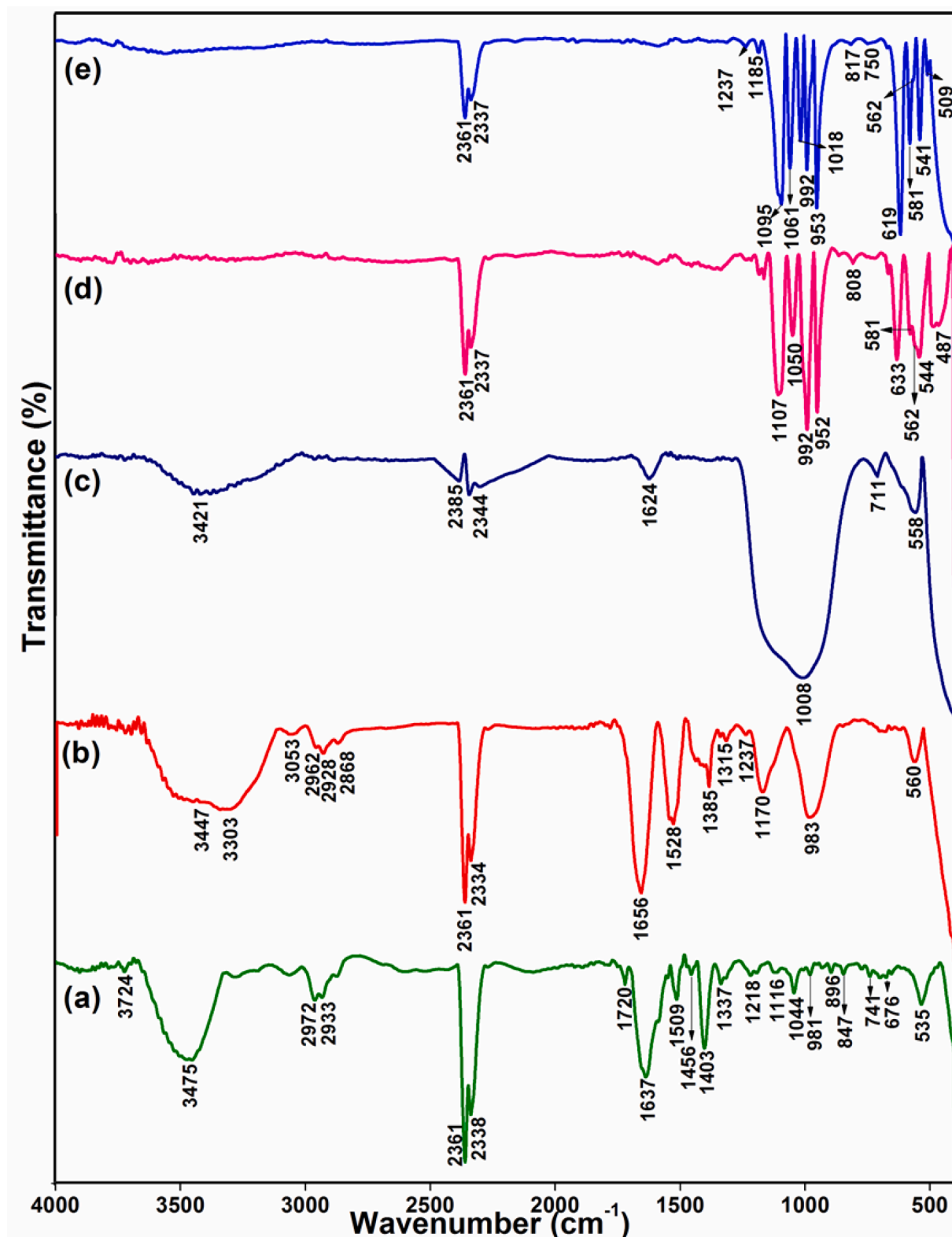


Fig. 4. FT-IR spectra of (a) *P. semisulcatus* extract, (b) CeNC@T1, (c) CeNC@T2, (d) CeNC@T3 and (e) Ag-CeNC@T3.

2.4. Biosynthesis of Ag@CeO₂-CePO₄ NCs

Ag@CeO₂-CePO₄ NCs were synthesized adapting two methods. In the first method, AgNO₃ (0.08494 g) was dissolved in the prepared yellow gel (section 2.3) and exposed to sunlight while stirring magnetically for another 2 h. In an alternate method, an equal volume of

P. semisulcatus extract was added to a 0.005 M solution of AgNO₃ and left in sunlight for 1 h. AgNPs thus prepared were added to the above yellow-coloured gel and magnetically agitated for another 2 h. The prepared gel was then centrifuged and washed with plenty of water and air dried. Finally, it was ground and further annealed at 800 °C for 2 h and labelled as Ag-CeNC@T3. The second method was followed owing to the more

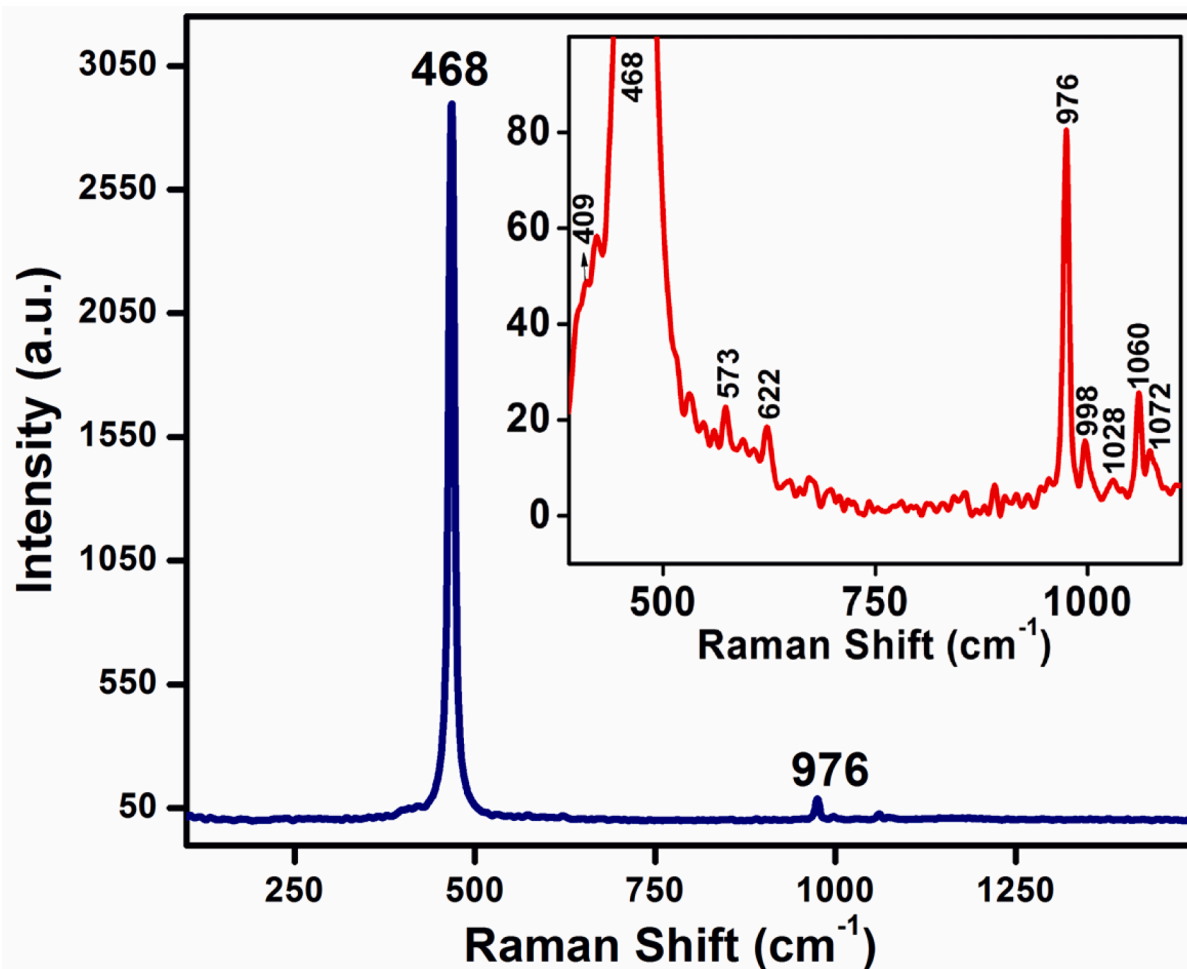


Fig. 5. Raman spectrum of CeNC@T3. Inset: Expanded view from 1110 cm^{-1} to 389 cm^{-1} .

intense colour of AgNPs.

2.5. Biosynthesis of CePO_4 NC

CePO_4 NC was prepared by dissolving 2.1661 g of cerium nitrate hexahydrate ($\text{Ce}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$) in 100 mL of water (0.05 M) and 100 mL of the *P. semisulcatus* extract was added dropwise slowly (1:1 v/v) while stirring magnetically. The reaction mixture was further subjected to heating at 60 °C for 1 h. This resulted in the formation of a white gel which was further stirred for 2 h using a magnetic stirrer at 1500 rpm. The resultant CePO_4 NC was further centrifuged, washed several times with water and subjected to air drying ($\text{CePO}_4\text{NC@T1}$).

A schematic representation for the synthesis of different NCs mediated by *P. semisulcatus* extract is presented in Scheme 1.

2.6. Characterization

UV–Visible Diffuse Reflectance Spectra (UV–Vis DRS) of synthesized NCs were captured using a UV–Visible spectrophotometer (JASCO, V-650 model) between 200 and 900 nm, and emission studies were carried out using a JASCO FP-8300 spectrofluorometer. A Shimadzu 8400S model instrument was used to record the Fourier Transform Infrared (FT-IR) spectra of dried shrimp extract and NCs in the range of 4000–400 cm^{-1} in KBr pellets. The crystallinity of the NCs was investigated using X-ray Diffraction (XRD) spectra obtained from an X'Pert Pro-PAnalytic Diffractometer with $\text{CuK}\alpha$ radiations (30 mA, $\lambda = 1.5406 \text{ \AA}$ and 40 kV). Raman analysis was performed using a Horiba XploRA PLUS Raman microscope with laser wavelength 638 nm, and Energy

Dispersive X-ray (EDX) analysis of NCs was done using Bruker Nano, GmbH, D-12489 instrument with an accelerating voltage of 30 kV. X-ray Photoelectron Spectroscopy (XPS) analysis was carried out to identify the chemical state of each element using PHI VersaProbe III instrument, which used a monochromatic Al- $\text{K}\alpha$ X-ray source ($h\nu = 1486.6 \text{ eV}$) with constant analyzer energy (26 eV) and binding energy was calibrated to 284.6 eV, which corresponds to the C-1s atom. High Resolution Transmission Electron Microscopy (HR-TEM) and Selected Area Electron Diffraction (SAED) measurements of the NCs were performed on JEOL 2100F and JEOL JEM 2100 PLUS instruments with an accelerating potential of 200 kV. HR-TEM images were acquired by dispersing the sample in an alcoholic medium that was cast onto copper grids and dried at ambient temperature.

2.7. Sensing of metal ions

To assess the potential of NCs as fluorescent sensors, different metal ions were examined at neutral pH under ambient conditions. For this, 10 mg of the NCs were dispersed in 30 mL bidistilled water (0.33 mg/mL) by continuously stirring with a magnetic stirrer throughout the experiment. 1 mL was taken from the dispersion setup and mixed with equal volume of target metal ions solution, to a total volume of 2 mL (0.05 M). The mixture was shaken for one minute, and the fluorescence spectra were then obtained with excitation at 220 nm. The fluorescence intensity of bio-derived NCs was enhanced by Pb^{2+} , Hg^{2+} and Cd^{2+} and hence their potential to detect these metal ions was studied in detail in the concentration range of $5.0 \times 10^{-9} \text{ M}$ to $4.0 \times 10^{-5} \text{ M}$. The Limit of Detection (LOD) was calculated using the formula: $\text{LOD} = 3\sigma / \text{slope}$,

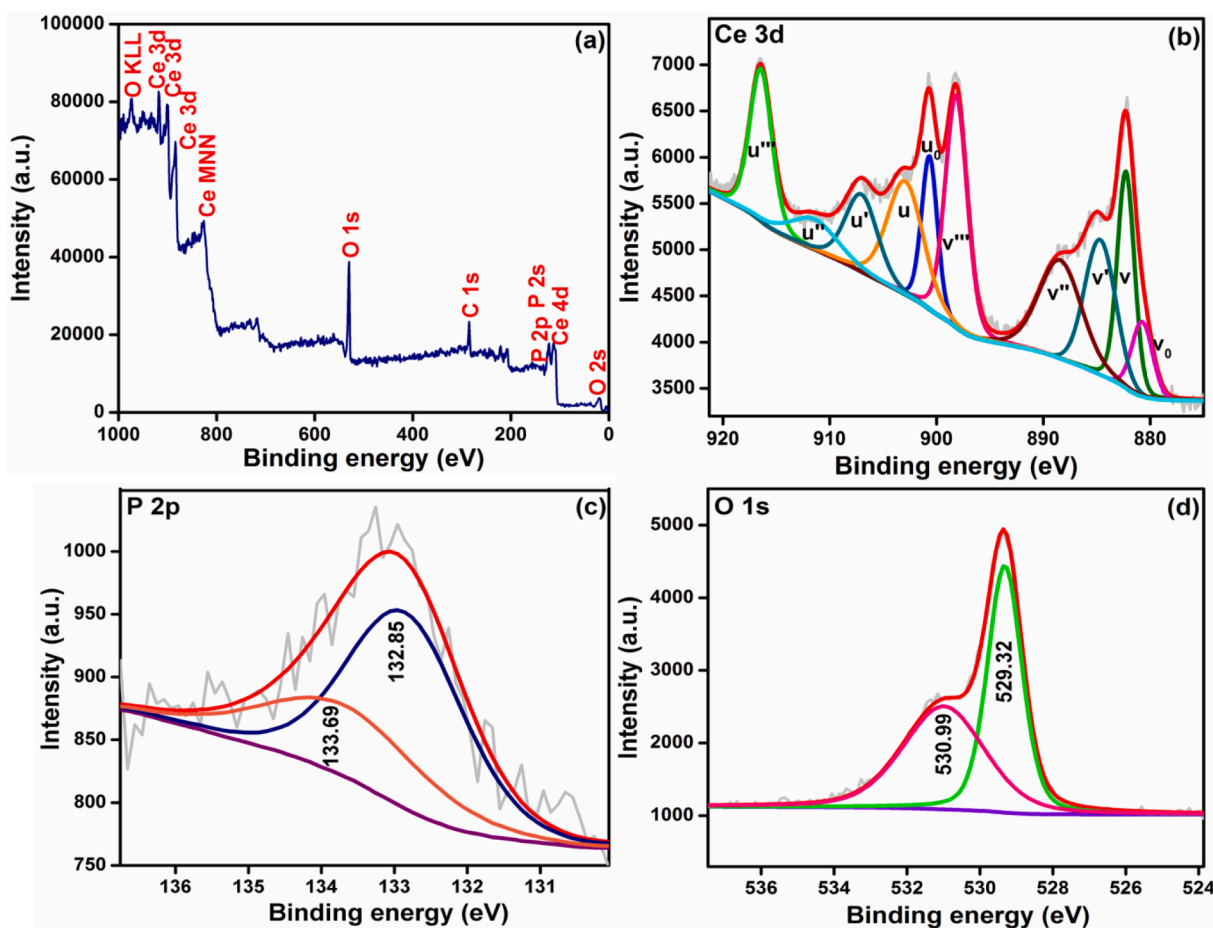


Fig. 6. XPS of CeNC@T3 (a) Survey spectrum, Core level spectra of (b) Ce 3d, (c) P 2p and (d) O 1 s.

where σ denotes the standard deviation of the blank ($n = 10$) (Gracia et al., 2022; Daphne et al., 2023; Gracia et al., 2023).

2.8. Evaluation of Sun protection factor (SPF) and cytotoxicity

SPF was detected using the method of Khan et al. (Khan, 2018) and the procedure is reproduced in S1 for quick reference. The *in vitro* screening of cytotoxicity of CeNC@T1 and CeNC@T2 was done by 3-(4,5-dimethylthiazol-2-yl)-2,5-diphenyltetrazolium bromide (MTT) assay (Gerlier and Thomasset, 1986) (S2). The cell viability was expressed using the formula:

$$\text{Cell viability}(\%) = \frac{\text{Average absorbance of treated cells}}{\text{Average absorbance of control}} \times 100$$

3. Results and discussion

P. semisulcatus has been reported to contain 36.9 % crude protein, 4.78 % lipid, 2.01 % carbohydrates and 56.43 % of water (El-Gendy et al., 2018) and minerals such as iron (0.51 %), phosphorus (0.16 %), potassium (0.43 %), magnesium (0.03 %), sodium (0.16 %), calcium (0.29 %), zinc (0.15 %), and copper (0.09 %) (AlFaris et al., 2022). The amino acids serine, threonine, tyrosine and histidine exist in phosphorylated form in *P. semisulcatus*. The phosphorus in the extract of *P. semisulcatus* has been estimated using Barton's reagent (Ishwarya et al., 2016); (Potel et al., 2018).

On adding the aqueous extract of *P. semisulcatus* to cerium (IV) solution, the protein is adsorbed to the CeO₂ gel formed (section 2.3) and reduces some of the Ce(IV) to Ce(III) which intrinsically facilitates oxygen vacancies. The phosphate ions in the shrimp extract stabilize part

of Ce(III) as CePO₄ (Fig. 1). The small size of CeO₂ (section 3.7) is the result of simultaneous formation of CePO₄ in the nanocomposite. The concentration of oxygen vacancies in ceria, reflected by the Ce³⁺/Ce⁴⁺ ratio, is influenced by various factors including nanoparticle size, morphology, dopants, surface area and the exposed high-energy facets of the ceria nanoparticles (Vinothkumar et al., 2018); (Noor et al., 2021).

3.1. X-ray diffraction analysis

XRD spectra of NCs delineated in Fig. 2. show typical peaks of CeO₂ and CePO₄ phases. The 2θ values of well-defined diffraction peaks along with crystallographic (*hkl*) planes are given in Table S1. The peaks reflect cubic fluorite structure of ceria (ICDD card no. 00-034-394) with $fm\bar{3}m$ space group (Vinothkumar et al., 2018). CePO₄ is expected to be in monoclinic phase at 550 °C. However, all the distinct peaks associated with monoclinic phase of CeO₂ are not observed in the diffraction pattern at this temperature owing to the convolution. Formation of CePO₄ on the surface of CeO₂ influences the particle size of nanoceria (Noor et al., 2021).

In CeNC@T1, the prophenoxidase peaks associated with the shrimp extract are located at 13.71 and 32.40 (Ishwarya et al., 2016), the peaks at 30.44 and 64.53 are due to lectin and the peak at 32.40 is attributed to peptide crustin (Sivakamavalli et al., 2020). Diffraction pattern of CeNC@T3 shows characteristic peaks corresponding to cubic CeO₂ NPs and monoclinic CePO₄ NPs and exhibits 66.56 % crystallinity. The peaks at 28.53, 33.06, 47.46, 56.33, 59.08, 69.42, 76.61 and 79.09 correspond to CeO₂ phases (ICDD card no. 00-034-394) (Vinothkumar et al., 2018). Monoclinic phase of CePO₄ is confirmed by diffraction peaks at 12.31, 21.44, 26.99, 42.02, 45.96 and 52.65 (JCPDS card

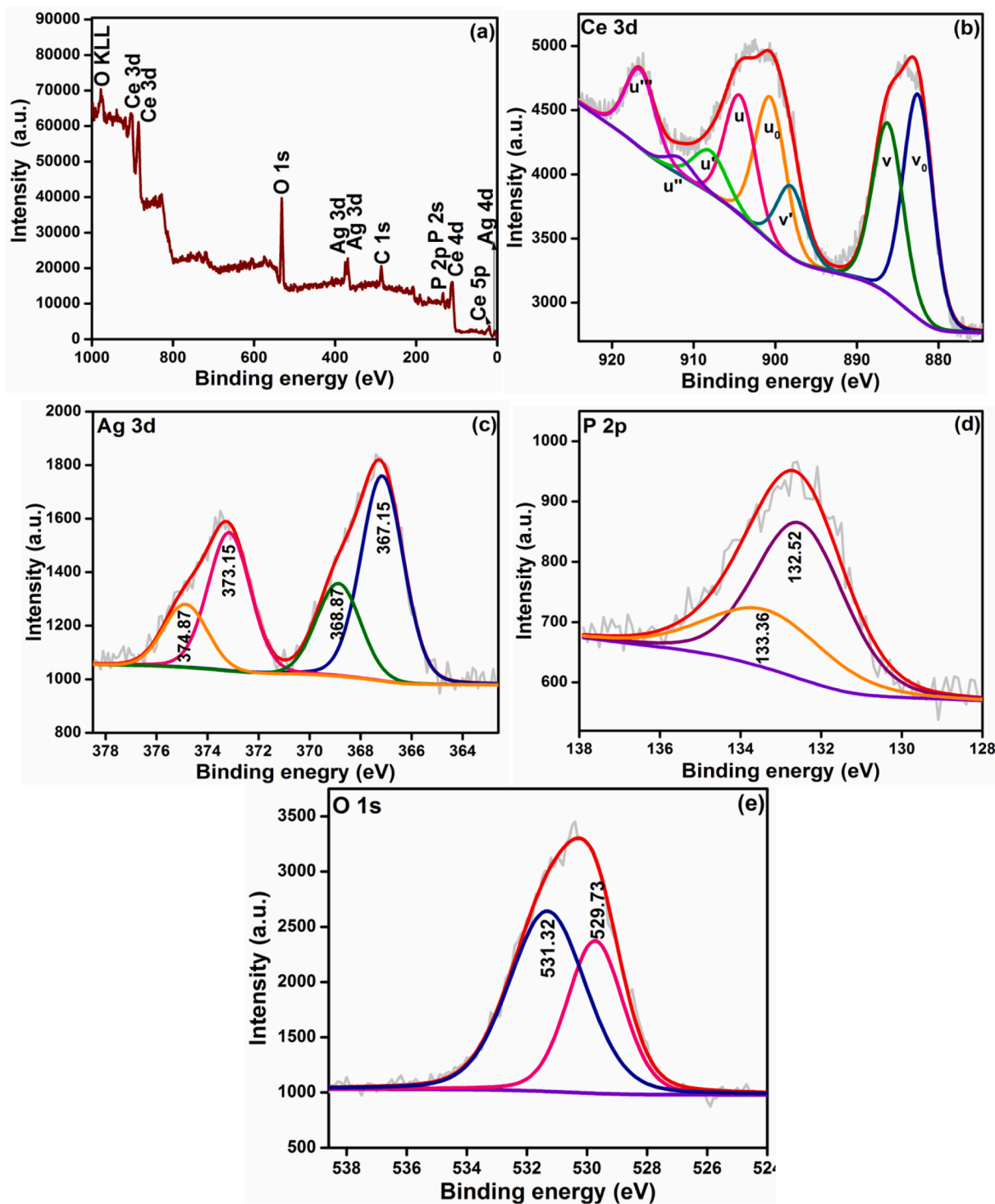


Fig. 7. XPS of Ag-CeNC@T3 (a) Survey spectrum, Core level spectra of (b) Ce 3d, (c) Ag⁺⁰, (d) P 2p and (e) O 1 s.

no.32–0199) (Noor et al., 2021).

Ag-CeNC@T3 shows characteristic diffraction peaks corresponding to CeO₂ at 28.55, 33.09, 47.50, 56.39, 59.19, 69.43, 76.33 and 79.17 which account for (111), (200), (220), (311), (222), (400), (331) and (420) planes respectively (ICDD card no. 00–034-394) (Vinothkumar et al., 2018). In Ag-CeNC@T3, peaks due to Ag⁺ ion are detected at 34.42 (Yang et al., 2016) and 41.09 (Shi et al., 2016). The peaks located at 12.03, 21.21, 26.95, 31.15, 42.22, 46.06, 48.56, 51.81 correspond to the (100), $\bar{1}$ (111), (120), (102), $\bar{1}$ (03), (212), (103) and $\bar{2}$ (32) planes of CePO₄ NPs (JCPDS card no.32–0199) (Verma and Bamzai,

2014).

A study of literature reveals that phosphorus in the extract of *P. semisulcatus* can serve as the precursor for the CePO₄ phase (Noor et al., 2021). The presence of CePO₄ phase has been confirmed by the XRD patterns. Additionally, EDX analysis further supports this finding by indicating that CeNC@T3 contains 4.52 % phosphorus (Fig. S1). The crystallite size of the NCs is ascertained using Debye-Scherrer's formula represented as $D = k\lambda/\beta \cos \theta$, where D, k, λ , β and θ represent the crystallite size, Scherrer constant (0.9), wavelength of X-ray beam, FWHM and the diffraction angle respectively. The average particle size

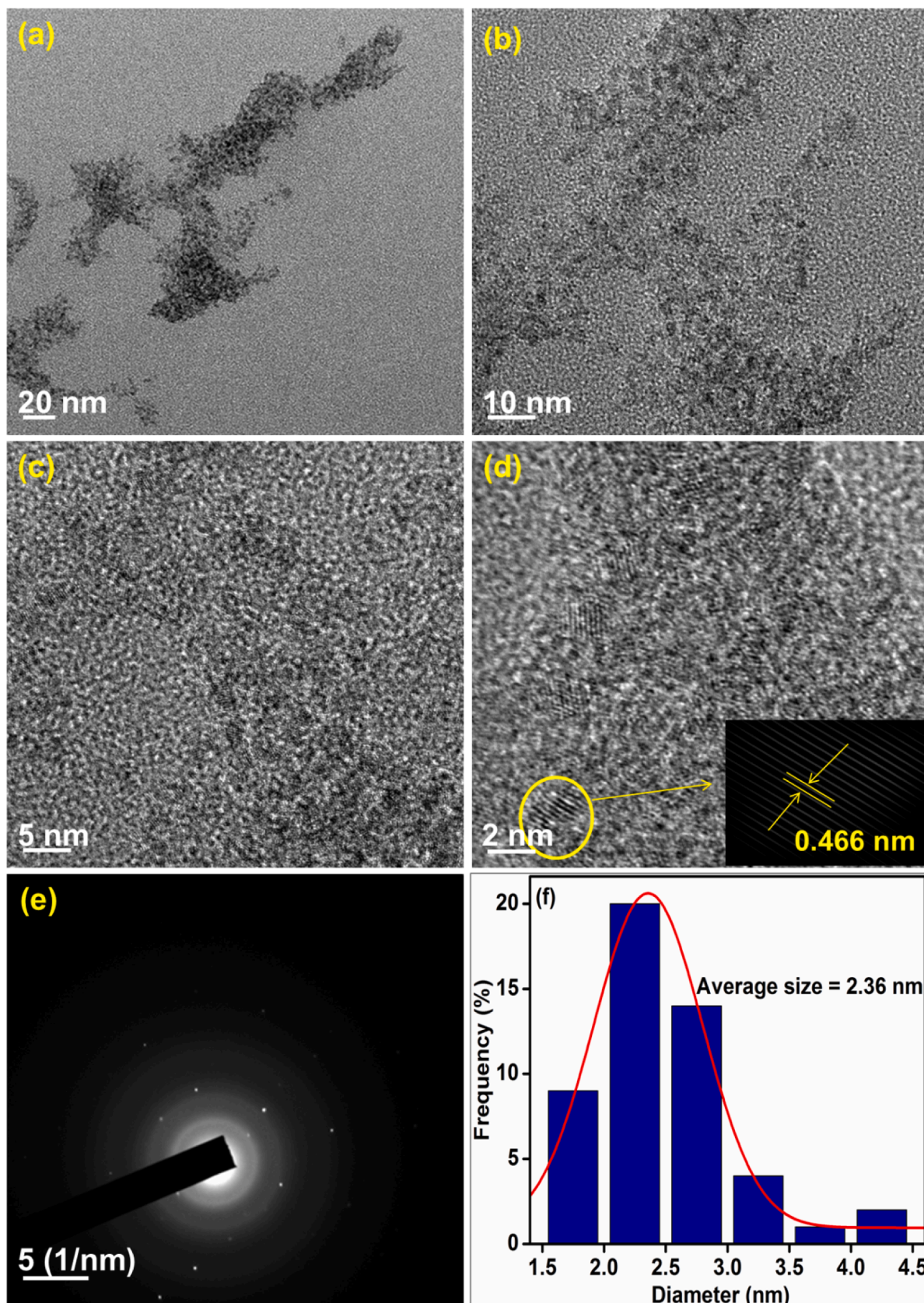


Fig. 8. HR-TEM analysis: (a)-(d) images of CeNC@T1 at different magnifications, (e) SAED pattern and (f) Histogram representing particle size distribution.

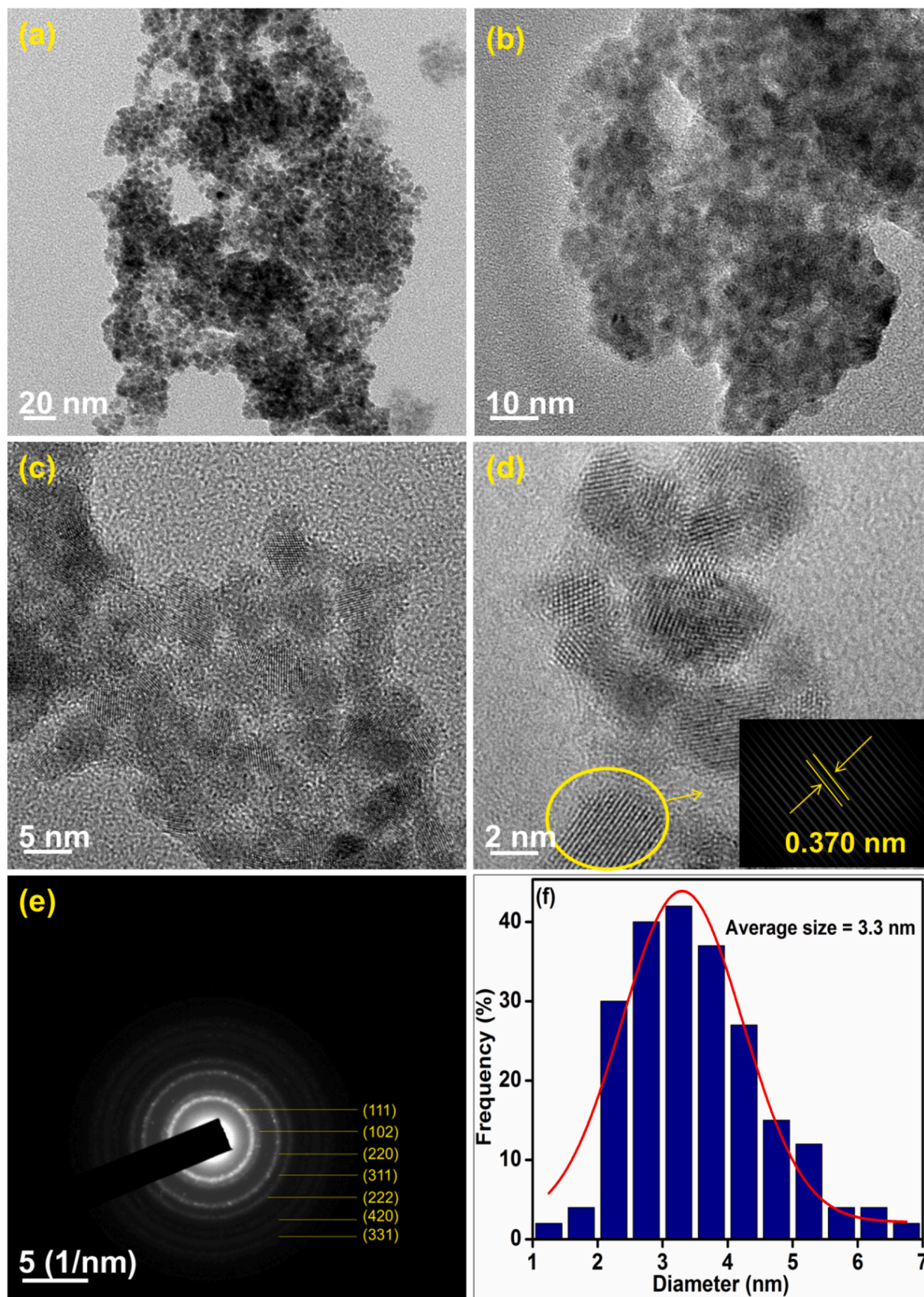


Fig. 9. HR-TEM analysis: (a)-(d) images of CeNC@T2 at different magnifications, (e) SAED pattern and (f) Histogram representing particle size distribution.

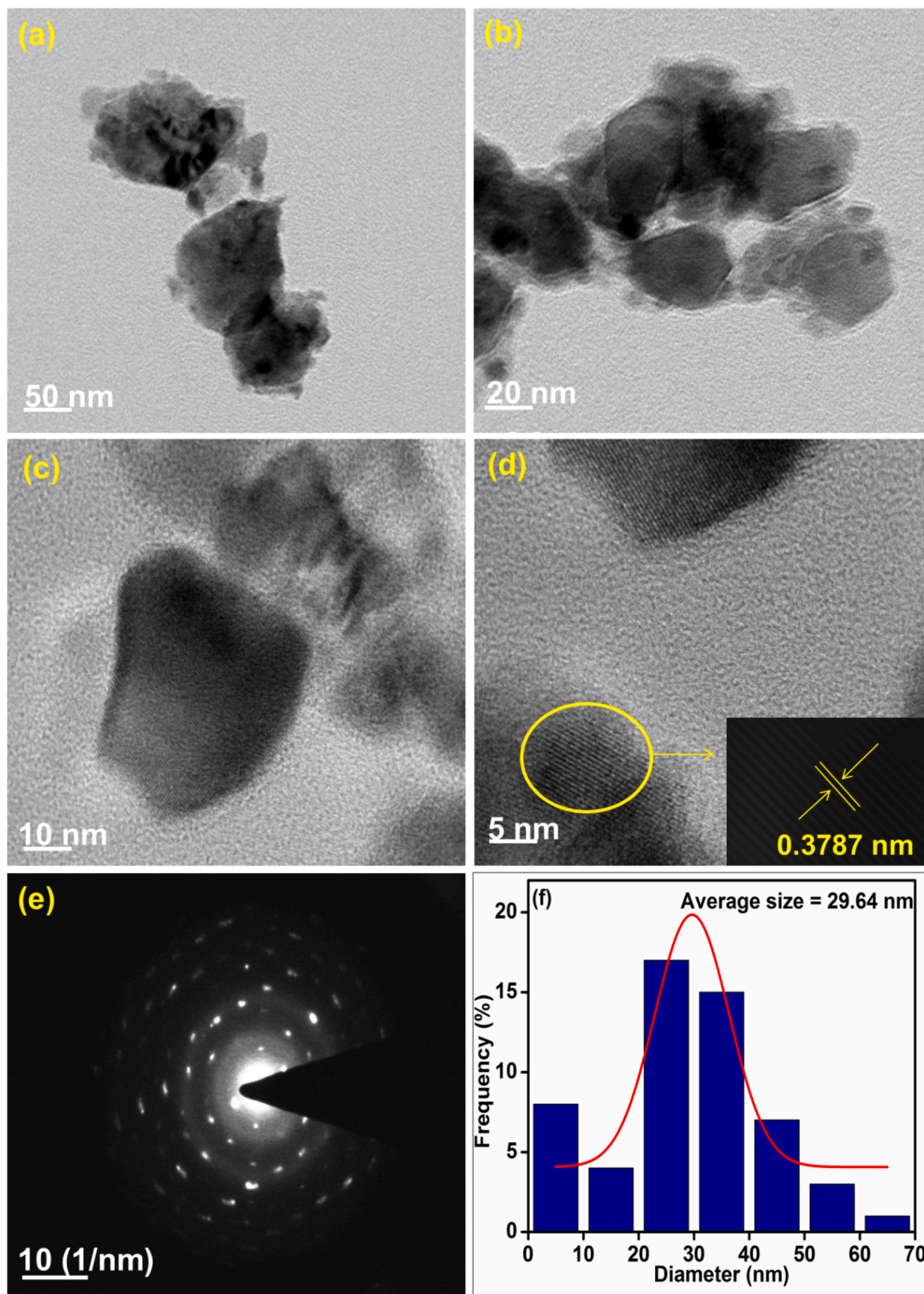


Fig. 10. HR-TEM analysis: (a)-(d) images of CeNC@T3 at different magnifications, (e) SAED pattern and (f) Histogram representing particle size distribution.

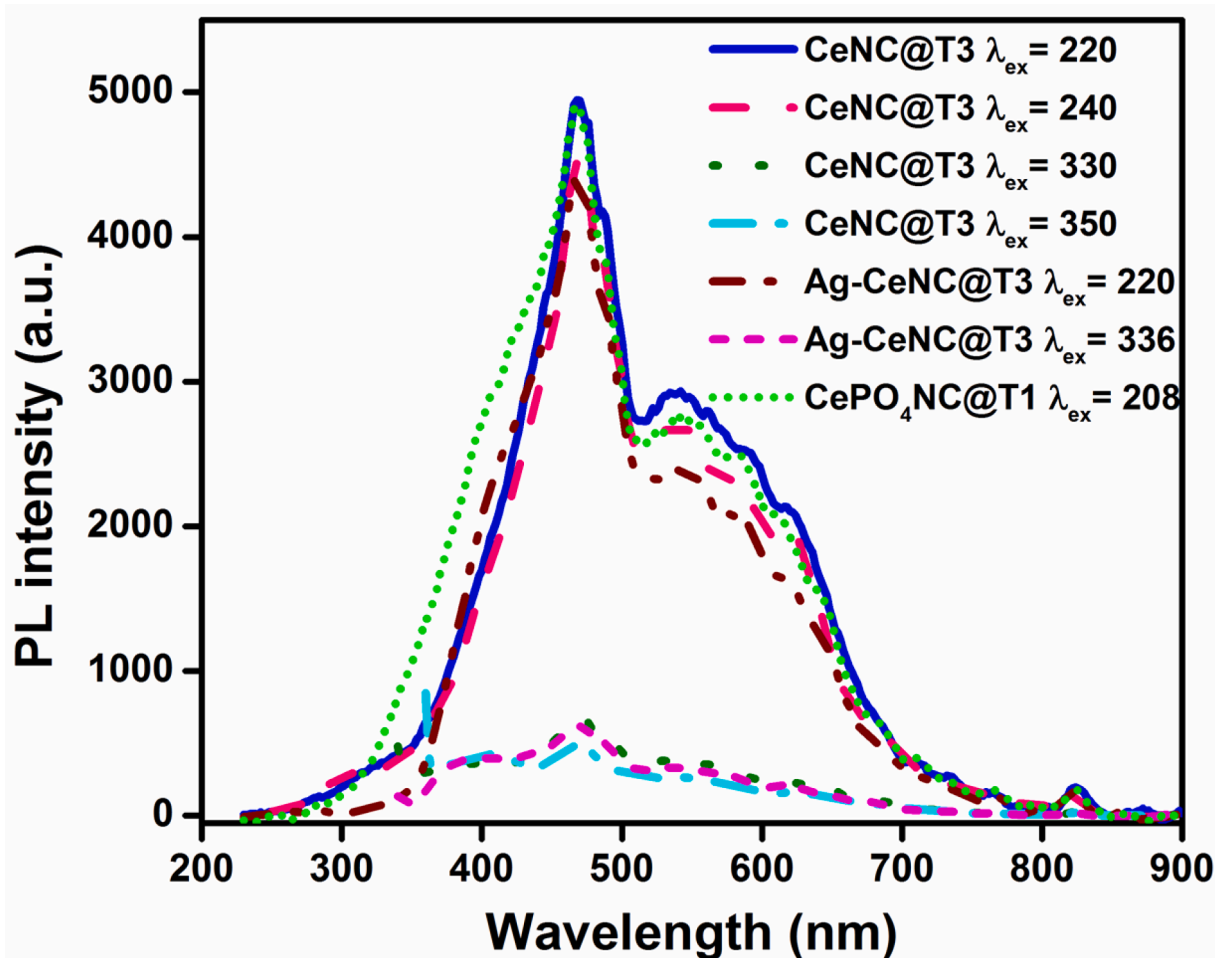


Fig. 11. Fluorescence spectra of solid CeNC@T3, Ag-CeNC@T3 and CePO₄NC@T1 at different excitation wavelengths.

of CeNC@T1, CeNC@T2, CeNC@T3 and Ag-CeNC@T3 has been found to be 4.78 nm, 6.51 nm, 44.29 nm and 32.73 nm respectively. With a marked increase in temperature from 550 °C to 800 °C, hexagonal phase of CePO₄ transforms to monoclinic phase, which is accompanied by an increase in size of the nanoparticles.

3.2. UV-Vis DRS analysis

UV-Vis DRS spectral analysis of *P. semisulcatus* derived NCs was studied between 200 and 900 nm, and the results are compiled in Fig. 3. CeNC@T1, CeNC@T2 and CeNC@T3 show broad peaks centered at 352, 356 and 352 nm, humps around 256, 250 and 256 nm respectively, and a broad tail extending up to 800 nm. A closer examination of the CeNC@T3 spectrum reveals the presence of additional humps at 212 and 304 nm. The absorption spectra of CeNC@T2 and CeNC@T3 show an almost identical absorption edge in the visible region exhibiting a small shift in λ_{max} . The difference observed in the slope of the absorption curve is attributed to the heightened degree of crystallinity in both the oxide and phosphate phases of cerium at elevated temperatures (Vinothkumar et al., 2018).

In all these spectra, the broad peaks around 355 nm are typical of CeO₂. Ce(IV) is a 4f⁰ system, where 4f-4f transition is not possible. The intense yellow colour of Ce(IV) compounds is due to charge transfer to metal (O²⁻(2p) to Ce⁴⁺(4f) in the 300–450 nm region (Choudhury et al., 2014). The peaks observed at 212, 256, 288 and 304 nm in CeNC@T3 are due to Ce(III) of CePO₄ and originate from ²F_{5/2} (GS) to the 5d (ES) (200–320 nm range) (Vinothkumar et al., 2018; Dezfuli et al., 2014).

For comparison, CePO₄ NC was prepared using Ce(III) precursor

(section 2.5) and air-dried (CePO₄NC@T1). It has a major peak at 286 nm and a small peak at 208 nm, revealing the presence of Ce³⁺. Most of the peaks due to Ce³⁺ are masked by LMCT in Ce⁴⁺ (Fang et al., 2003). The broad peak observed at 352 nm in CeNC@T3 is blue shifted at 336 nm in Ag-CeNC@T3 with a hump at 254 nm and cut-off wavelength of 600 nm. The hump around 300 nm due to Ce(III) is not observed in this spectrum.

The band gaps of synthesised CeNC@T1, T2, T3 and Ag-CeNC@T3 have been calculated by applying Tauc equation (Jubu et al., 2022). UV-Vis DRS spectra of the CeNC@T1, T2 & T3, Ag-CeNC@T3 and CePO₄NC@T1 have distinct absorption bands around 354 nm and the optical band gaps have been calculated as 2.77, 2.72, 2.85, 3.19 and 3.71 eV respectively (Fig. 1 inset). Kubelka-Munk function has also been used to determine the optical band gap of the prepared nanocomposites (Noor et al., 2021). The optical band gaps of CeNC@T1, CeNC@T2, CeNC@T3, Ag-CeNC@T3 and CePO₄NC@T1 using Kubelka-Munk function are 3.03, 2.95, 3.08, 3.34 and 3.82 eV. A band gap of 2.9–3.3 eV is observed for bulk CeO₂ and band gap of 3.2–3.6 eV for both direct and indirect band transitions. CePO₄ NPs have been reported to have a direct band gap of around 3.0 eV. The reduction in both direct and indirect band gaps in this study in comparison to the band gap values of pristine CeO₂ and CePO₄ nanoparticles is reported to be due to the synergy that exists between the two phases in the NCs. The same trend is observed in CeO₂-InVO₄ and CeO₂-BiVO₄ nanocomposites (Noor et al., 2021).

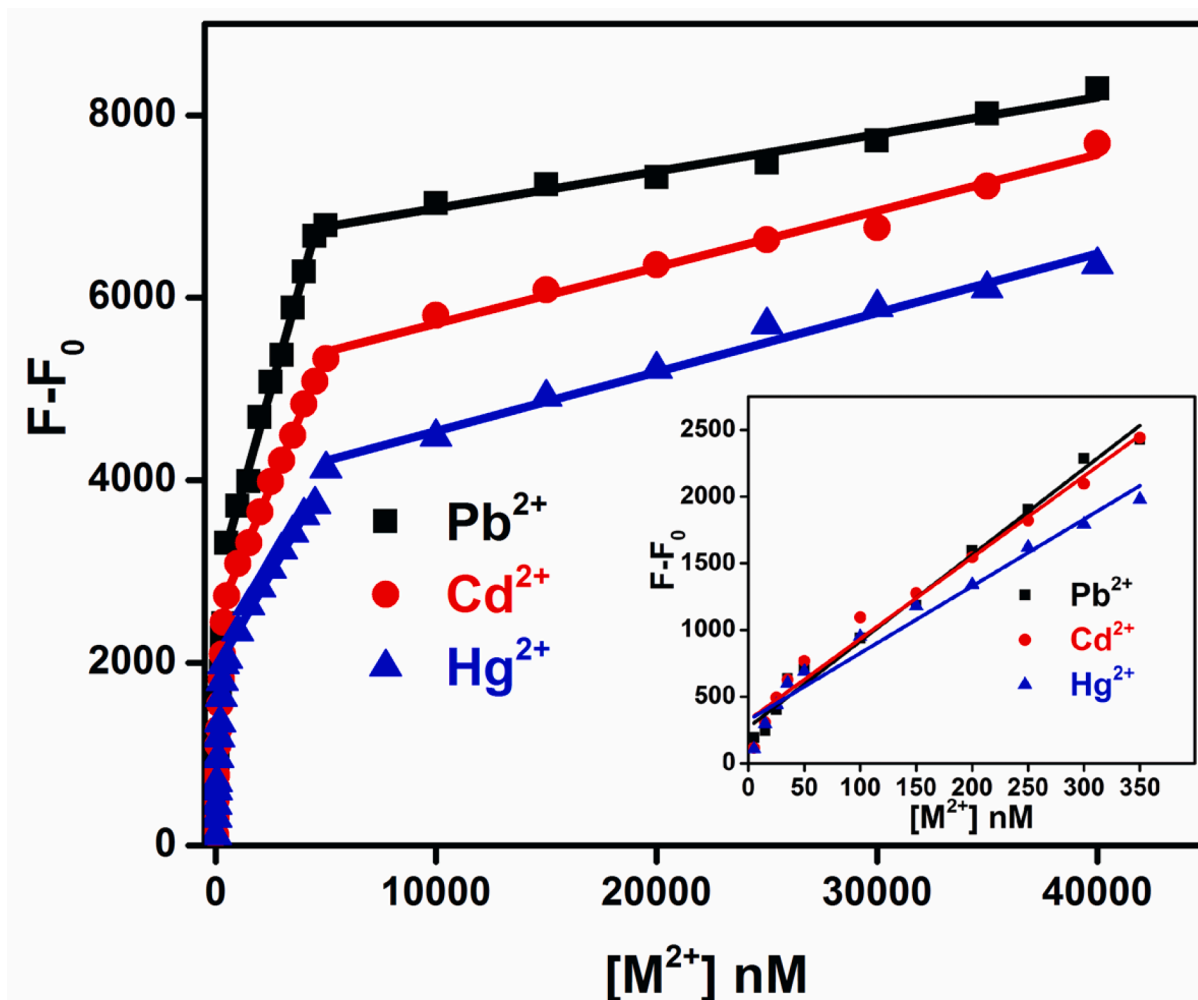


Fig. 12. Plot of fluorescence intensity ($F-F_0$) Vs $[M^{2+}]$, nM for CeNC@T3, Inset: Expanded view of 5 nM to 350 nM region.

3.3. FT-IR spectral analysis

The FT-IR spectral data of *P. semisulcatus* extract, CeNC@T1, CeNC@T2, CeNC@T3 and Ag-CeNC@T3 are delineated in Fig. 4. The FT-IR spectrum of the extract (Fig. 4a) displays a broad peak centered at 3475 cm^{-1} , which is attributed to $-\text{OH}/\text{NH}$ groups from the protein. The peaks at 2972 and 2933 cm^{-1} indicate $-\text{CH}$ stretching, while the band at 1637 cm^{-1} is attributed to the $\text{C}=\text{O}$ of the peptide group. The peak at 1044 cm^{-1} corresponds to $\text{C}-\text{O}$ stretch (Cozzolino and Murray, 2012). A specific feature of secondary amides is represented by the peak at 1456 cm^{-1} (Power and Cozzolino, 2020). The peak at 981 cm^{-1} confirms the presence of phosphorylated proteins (Jeyachandran et al., 2020), while the antisymmetric stretching of the $\text{C}-\text{O}-\text{C}$ skeleton occurs at 1116 cm^{-1} and the band at 2338 cm^{-1} , observed in all spectra, is indicative of trapped CO_2 (Cheng et al., 2021).

In the FT-IR data of CeNC@T1, $\text{Ce}-\text{O}$ stretching is at 560 cm^{-1} (Patil et al., 2016). The broad overlapping bands at 3447 cm^{-1} and 3303 cm^{-1} are due to $-\text{NH}/\text{OH}$ of the protein or traces of H_2O molecules. The peak at 1656 cm^{-1} is attributed to $\text{C}=\text{O}$ of the peptide group. The peak located at 983 cm^{-1} confirms $\text{P}-\text{O}$ of distorted PO_4^{3-} tetrahedra. The FT-IR spectrum of CeNC@T2 shows a broad band centered around 3421 cm^{-1} and a peak at 1624 cm^{-1} corresponding to the stretching and bending modes of H_2O . The presence of bands due to H_2O molecules even after drying at $550\text{ }^\circ\text{C}$ indicates that CePO_4 is crystallized in hexagonal form and contains water of hydration, $\text{CePO}_4 \cdot n\text{H}_2\text{O}$ (Lucas et al., 2004). The formation of CeO_2 nanoparticles is corroborated by the $\text{Ce}-\text{O}$ stretching observed at 558 cm^{-1} (Patil et al., 2016). The broad intense band at

1008 cm^{-1} and the peak observed at 711 cm^{-1} represent PO_4^{3-} groups (Parimi et al., 2019).

FT-IR studies of CeNC@T3 show $\text{Ce}-\text{O}$ stretching at 562 cm^{-1} and the peak due to PO_4^{3-} (broad signal located at 1008 cm^{-1} in the case of CeNC@T2, is structured into many peaks. The peaks at 952 , 992 , 1050 , 1107 cm^{-1} correspond to $\text{P}-\text{O}$ stretching modes of PO_4^{3-} and the signals at 544 , 583 , 633 cm^{-1} are due to bending vibrations of $\text{O}-\text{P}-\text{O}$ and $\text{O}=\text{P}-\text{O}$, which confirm the presence of distorted tetrahedron of PO_4^{3-} in monoclinic phase (Bouddouch et al., 2021; Cheng et al., 2021). The peaks associated with H_2O molecules (3421 cm^{-1} and 1624 cm^{-1}) observed for CeNC@T2 disappear in the spectrum of CeNC@T3 (dried at $800\text{ }^\circ\text{C}$) signifying the formation of monoclinic CePO_4 from hexagonal $\text{CePO}_4 \cdot n\text{H}_2\text{O}$.

Almost similar peaks are identified in the FT-IR profile of Ag-CeNC@T3. $\text{Ce}-\text{O}$ stretching is observed as a minor peak at 562 cm^{-1} which is masked by strong $\text{P}-\text{O}$ bending vibrations. The $\text{P}-\text{O}$ stretching modes of distorted tetrahedra are observed at 953 cm^{-1} , 992 cm^{-1} , 1061 cm^{-1} and 1095 cm^{-1} . $\text{O}-\text{P}-\text{O}$ bending and $\text{O}-\text{P}$ bending corresponding to CePO_4 nanoparticles are located at 541 , 581 and 619 cm^{-1} (Bouddouch et al., 2021).

3.4. Raman spectral analysis

Raman spectrum of CeNC@T3 (annealed at $800\text{ }^\circ\text{C}$) was recorded from 3000 cm^{-1} to 450 cm^{-1} and the results are presented in Fig. 5. The inset represents a section of Raman spectrum in the range 1110 cm^{-1} to 389 cm^{-1} . The most intense signal observed at 468 cm^{-1} is associated

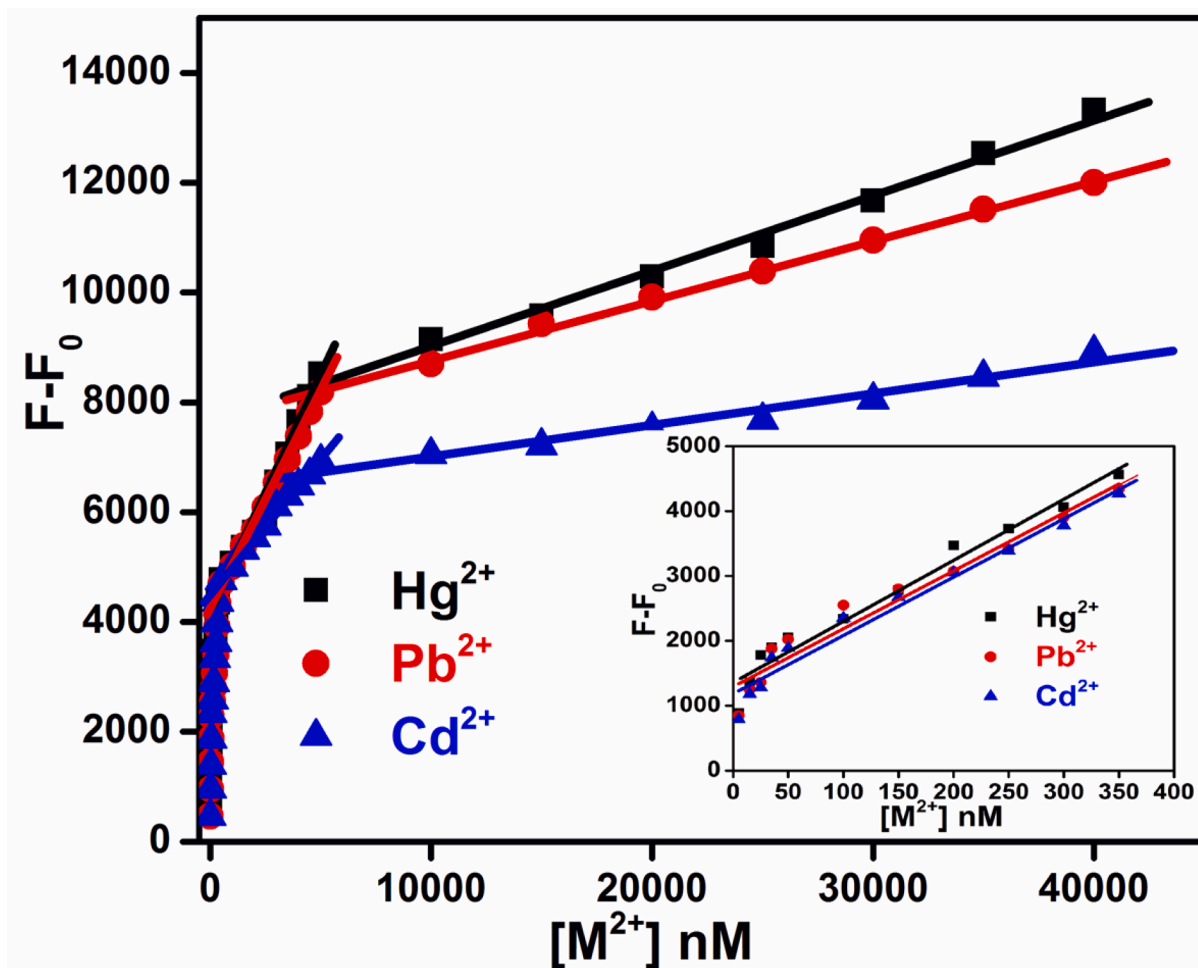


Fig. 13. Plot of fluorescence intensity ($F-F_0$) Vs $[M^{2+}]$, nM for Ag-CeNC@T3, Inset: Expanded view of 5 nM to 350 nM region.

with the F_{2g} band of CeO_2 (Aslam et al., 2016). A slight shift from 465 cm^{-1} corresponding to single crystal ceria is related to the polycrystallinity of the nanocomposite (Schilling et al., 2017) or to the presence of defects, phonon confinement and strain effects. The broad nature of this peak indicates the smaller crystallite size of CeO_2 NPs, the presence of defects and strain effects (Gouadec and Colombar, 2007).

Combined $E_g + A_{1g}$ stretching modes are observed as a small hump at 409 cm^{-1} (Schilling et al., 2017). The intense peak observed at 976 cm^{-1} ascertains PO_4^{3-} tetrahedron present in the nanocomposite. The minor peaks at $998, 1028, 1060, 1072\text{ cm}^{-1}$ are also due to monoclinic $CePO_4$ (Allen et al., 2021). The peak at 1072 cm^{-1} is associated with anti-symmetric stretching of the P-O bond (Vinothkumar et al., 2018). The weak bands at 573 cm^{-1} and 622 cm^{-1} confirm the antisymmetric bending modes of PO_4^{3-} group (Vinothkumar et al., 2018; Cheng et al., 2021). The bands of weak intensity resulting from the different phonon symmetry modes corresponding to second order scattering are also observed. Oxygen vacancies present in the cubic fluorite structure are further confirmed by the signal at 1060 cm^{-1} (Jayakumar et al., 2019). The 'D band' corresponding to the defects in nanocomposite is observed at 573 cm^{-1} (Calvache et al., 2019).

3.5. X-ray photoelectron spectral analysis

The chemical states of Ce and O and the Ce^{3+}/Ce^{4+} ratio at the surface were assessed using XPS between 0 and 1000 eV. The deconvoluted 3d XPS spectrum of CeNC@T3 is given in Fig. 6b, which consists of 6 doublets (spin-orbit) (Vinothkumar et al., 2018). The peaks positioned at 882.27, 888.43, 898.16, 902.92, 911.39 and 916.46 eV are

assigned to v, v', v'', u, u' and u'' of $3d_{5/2}$ and $3d_{3/2}$ states Ce^{4+} respectively. In addition, it shows four components at 880.75 (v_o), 884.68 (v'), 900.67 (u_o) and 907.03 eV (u') associated with the main and satellite peaks of Ce^{3+} state of $CePO_4$ (Beche et al., 2008; Konyshva and Francis, 2013). $Ce(IV)$ present in cerium compounds is characterized by the satellite peak (u) associated with $Ce\ 3d_{3/2}$. The highest binding energy peaks, u and v respectively are at 916.46 ($3d_{3/2}$) and 898.16 eV ($3d_{5/2}$) and the spin-orbit splitting is 18.3 eV (Matussin et al., 2023). The $[Ce^{3+}]$ and $[Ce^{4+}]$ in CeNC@T3 are 31.7 % and 68.31 %, respectively (Vinothkumar et al., 2018).

P 2p core level spectrum has an asymmetric peak centered at 133.1 eV which is deconvoluted further into two peaks identified at 132.85 eV (P $2p_{3/2}$) and 133.69 eV (P $2p_{1/2}$) (Fig. 5c), representing phosphorus in $CePO_4$ (Vinothkumar et al., 2018; Fu et al., 2022). The O 1s core level spectrum of CeNC@T3 shown in Fig. 5d has two peaks at binding energies 529.32 and 530.99 eV. The presence of oxygen in Ce-O lattice is confirmed by the prominent peak at 529.32 eV. P-O bonding from PO_4^{3-} octahedron corresponding to higher binding energy is located at 530.99 eV (Vinothkumar et al., 2018; Vinothkumar et al., 2018).

Chemical bonding nature of Ag-CeNC@T3 has also been analyzed applying XPS technique. The survey spectrum in Fig. 7a demonstrates that the sample comprises of Ce, O and Ag. The deconvoluted Ce 3d XPS spectrum of Ag-CeNC@T3 has peaks at 882.53, 886.22, 898.08, 900.63, 904.32, 907.95, 911.65 and 916.56 eV (Chen et al., 2022). The concentration of Ce^{3+} and Ce^{4+} species in Ag-CeNC@T3 are found to be 22.44 % and 77.56 %, respectively. One notable feature of this spectrum is that the peaks corresponding to Ce^{3+} ions are not located. This is in agreement with DRS data (section 3.1) in which LMCT peak of

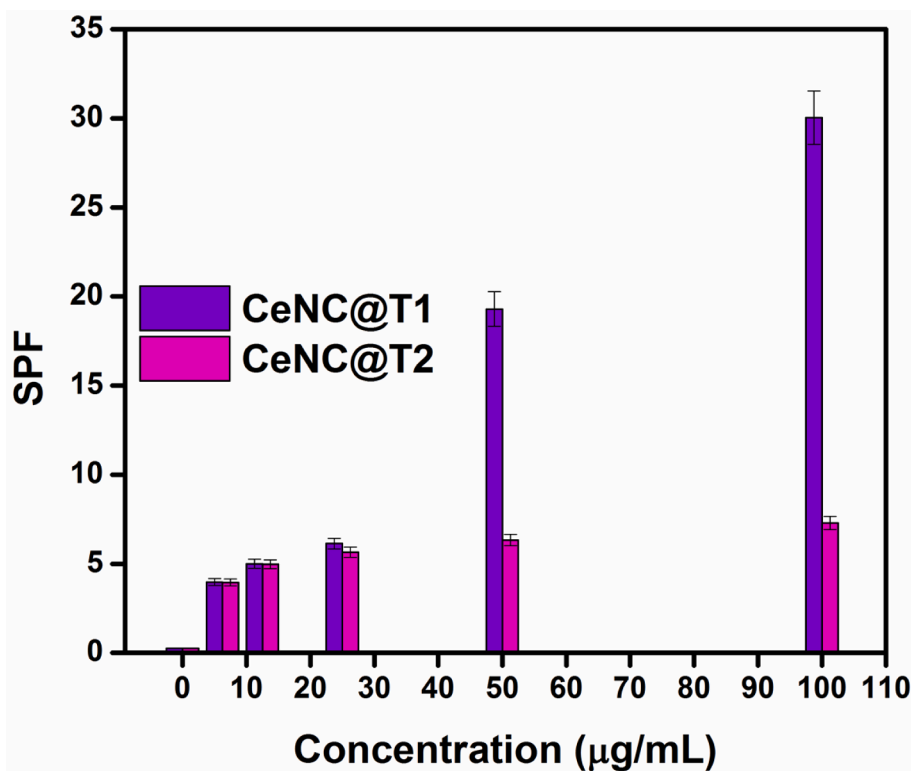


Fig. 14. SPF of CeNC@T1 and CeNC@T2.

Table 1

Comparison of reported SPF values of NMs with the value of CeNC@T1.

Nanomaterials (NMs)	Source	Dose of NMs	SPF value	Reference
5 % Ni doped CeO ₂ NPs	<i>Salvadora persica</i>	10,000 µg/mL	48.52	(Miri et al., 2020)
5 % Ni doped CeO ₂ NPs	<i>Stevia rebaudiana</i>	10,000 µg/mL	43.47	(Khatami et al., 2019)
CeO ₂ NPs	<i>Ziziphus jujube</i>	10,000 µg/mL	41.52	(Miri et al., 2020)
CeO ₂ NPs	<i>Musa sapientum</i>	10,000 µg/mL	38.76	(Miri et al., 2021)
CeNC@T1	<i>Penaues semisulcatus</i>	100 µg/mL	30.04	This work

CeNC@T3 observed at 352 nm is shifted to 336 nm with narrowing of the peak and the band gap is increased to 3.19 eV from 2.85 eV. All these data suggest that on adding AgNPs to CeNC@T3, the defects associated with Ce³⁺ are reduced and some of the Ag⁰ are oxidized to Ag⁺ owing to the smaller work function of Ag compared to CeO₂ (4.2 Vs 4.7 eV) (Matussin et al., 2023; Chen et al., 2022) and both species coexist on the surface of Ag-CeNC@T3. All these data suggest that surface oxygen vacancies of ceria are affected by silver loading and silver predominantly exists as Ag⁰ and oxidation to Ag⁺ occurs to a smaller extent (Chen et al., 2022; Qu et al., 2013).

The deconvoluted 3d XPS spectrum of Ag has peaks at 367.15 eV due to Ag⁺ and the peaks at 368.87 eV, 373.15 eV and 374.87 eV correspond to Ag⁰. The peaks located at 368.87 and 374.87 eV correspond to Ag 3d_{5/2} and Ag 3d_{3/2} of metallic silver with spin-orbit coupling of 6 eV (Murugadoss et al., 2021). Further, in the case of Ag-CeNC@T3, the binding energy of Ag 3d_{5/2} shows a relative increase compared to that of pure metallic Ag (~368.2 eV) signaling the loading of AgNPs to the surface of CeO₂. This confirms that a new Fermi energy level is formed in Ag-CeNC@T3, and there is an interaction between Ag and CeO₂ in the nanocomposite (Firet et al., 2019).

In addition, the P 2p peaks are shifted to lower energies 132.52 eV (132.85 eV) and 133.36 eV (133.69 eV), in comparison with CeNC@T3. The deconvoluted O 1s XPS spectrum of Ag-CeNC@T3 has two peaks. While comparing with CeNC@T3, a slight shift to higher energy is observed, 529.73 eV (529.32 eV), 531.32 eV (530.99 eV). The peak at 529.73 eV is due to O₂ in the CeO₂ lattice, whereas the peak observed at 531.32 eV may be correlated O/OH of oxygen adsorbed on the surface or P-O bonding from PO₄³⁻ (Vinothkumar et al., 2018; Vinothkumar et al., 2018; Qu et al., 2013). No observable OH⁻ peaks in the CeO₂ nanoparticles adorned with Ag (Murugadoss et al., 2021).

3.6. EDX analysis

The EDX spectra of NCs is summarized in S3 and it features the signatures of cerium, oxygen, phosphorus and silver. In all NCs, a major peak is observed at 4.82 keV corresponding to cerium, and other minor peaks observed around 0.67, 0.87, 5.26, 5.63 and 6.01 keV are also attributed to cerium. In all the NCs, the presence of oxygen is confirmed from the prominent peak obtained at 0.52 keV. Similarly, a significant peak at 2.02 keV is attributed to the presence of phosphorus in all the four NCs (Vinothkumar et al., 2018). Additionally, in Ag-CeNC@T3, silver is confirmed by a salient peak observed at 2.96 keV along with minor peaks at 0.36 and 3.15 keV (Devi et al., 2022).

3.7. HR-TEM analysis

The surface morphologies of the synthesized nanocomposites were analyzed using HR-TEM and the representative images are presented in Figs. 8–10. The particles are found to be spherical and the SAED patterns reveal that the nanocomposites are highly crystalline. The average particle size of CeNC@T1, CeNC@T2 and CeNC@T3 are 2.36 nm, 3.30 nm and 29.64 nm correspondingly. The SAED pattern of CeNC@T2 has (1 1 1), (2 2 0), (3 1 1), (2 2 2), (4 2 0) and (3 3 1) planes of CeO₂ and (1 0 2) planes corresponding to the monoclinic phase of CePO₄. The crystalline nature of the nanocomposites is confirmed by the lattice fringes, and the

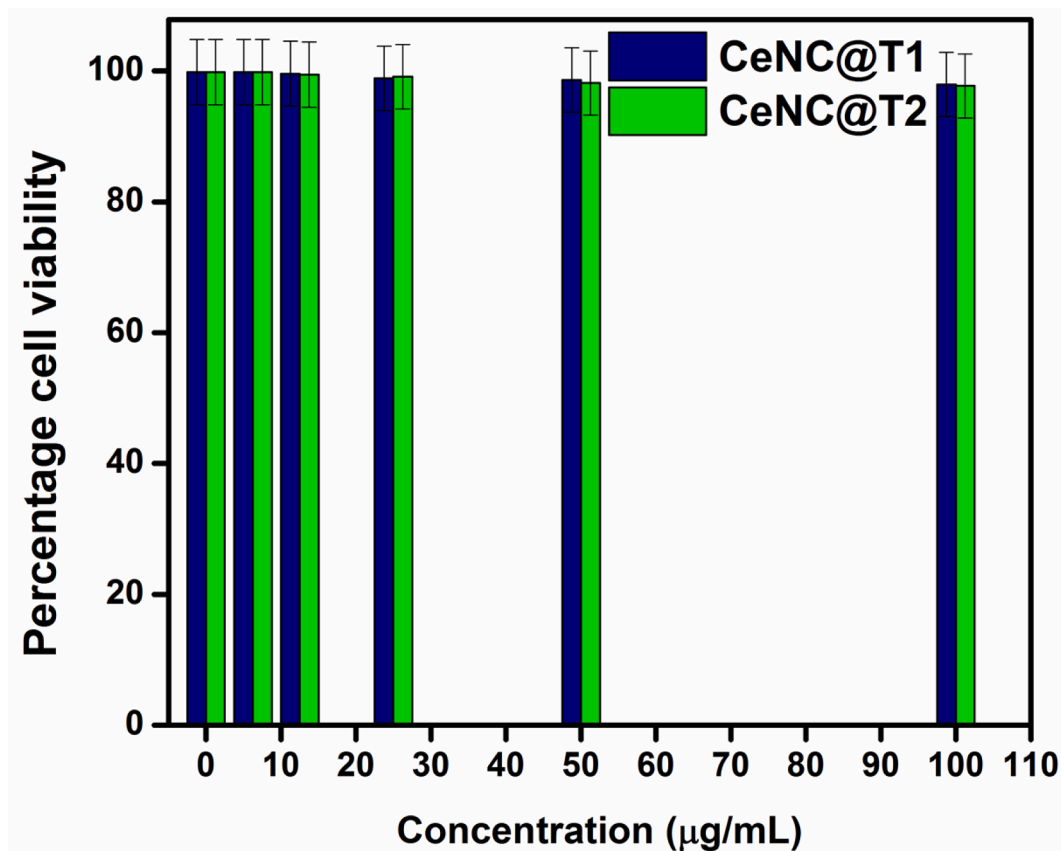


Fig. 15. Cytotoxic findings for CeNC@T1 and CeNC@T2.

SAED pattern, showing a ring-like pattern, further proves that the nanocomposites are crystalline, which corroborates with XRD data.

3.8. Fluorescence study

The colour of fluorescence emission of CeNC@T3 and Ag-CeNC@T3 are found out applying Commission Internationale de l'Eclairage (CIE) chromaticity diagram. CeNC@T3 emitted almost a white light and Ag-CeNC@T3 emitted light close to the blue light region, which are shown in Fig. S2 a & b. The room-temperature fluorescence spectra of CeNC@T3, Ag-CeNC@T3 and CePO₄NC@T1 were recorded as solid. The solid fluorescence spectra of NCs at different excitation wavelengths are given in Fig. 11.

CeO₂ fluoresces in UV region due to emission from 4f energy levels (Krishnan et al., 2013) but owing to the defects attributed to Ce³⁺ it can emit light in visible region (Mochizuki and Fujishiro, 2009). The 220 nm and 240 nm excited emission spectrum of CeNC@T3 has peaks at 470, 486, 542, 616 and 824 nm and similar less intensity emission peaks are observed in 330 nm and 350 nm excitations. Ag-CeNC@T3 nanocomposite has major peaks at 396, 468, 486, 538 and 618 nm. The highest intensity emission signal observed around 470 nm in all the spectra is due to emission from 4f energy level. The less intense bands peaking at 542 and 616 nm could be attributed to emissions originating from oxygen defects. The emission bands spanning from 400 to 500 nm in the CeO₂ samples are ascribed to transitions between various defect levels and the valence band (Aslam et al., 2016). The 5 excitation peaks corresponding to the 4f → 5d transition, induced by Ce³⁺ ions, and the peak at 300 nm is one of them (Qin and Niu, 2016). The emission spectrum of CePO₄ nanocomposite formed using Ce³⁺ precursor (section 2.5) is compared with those of CeNC@T3 and Ag-CeNC@T3. CePO₄NC@T1 (air dried) contains only CePO₄ in hexagonal structure. The emission peaks are almost the same, suggesting that the emission is

mainly due to CePO₄.

4. Applications

4.1. Sensing of Pb²⁺, Cd²⁺ and Hg²⁺

As the CeO₂-CePO₄ NCs display an intense fluorescence emission, their application to sense various metal ions is explored. The presence of various metal ions, especially Pb²⁺, Cd²⁺ and Hg²⁺ leads to an augmentation in the emission intensity (Fig. S3 & S4). Hence the potential of aqueous dispersions of CeNC@T3 to sense these heavy metals in the wide range of concentrations (10⁻⁹ M to 10⁻⁵ M) is investigated.

On mixing Pb²⁺ ions to the aqueous dispersion of CeNC@T3, the emission intensity proportionally enhanced with [Pb²⁺] (Fig. S5). When altering the concentration of Pb²⁺ from 5.0x10⁻⁹ M to 4.0x10⁻⁵ M, the fitting plot exhibits three clearly defined linear ranges (Fig. 12). The corresponding linear fitting equations are Y₁ = 6.4675X + 269.7334; Y₂ = 0.8546X + 2863.75; Y₃ = 0.04052X + 6574.321 with a correlation coefficient of 0.9934, 0.9978 and 0.9895 respectively, where X is the concentration of Pb²⁺ in nM. The LOD is found to be 2.41x10⁻⁸ M, 1.83x10⁻⁷ M, 3.85 x10⁻⁶ M in three regions respectively. The inset of Fig. 12 illustrates the magnified view of the low concentration range of Pb²⁺. This linear region is highly sensitive to the concentration of analytes. Hence it is worth developing CeNC@T3 as an inorganic fluorescent probe which can be easily prepared and is compatible with green chemistry for sensing heavy metals.

Similarly, the Cd²⁺ and Hg²⁺ ions also increase the emission intensity to a greater extent (Fig. S6 & S7). In these cases also, there are three linear regions in which the emission intensity is directly proportional while increasing the concentration of these ions 10,000 folds between 5.0x10⁻⁹ M to 4.0x10⁻⁵ M (Fig. 12). The regression equations for Cd²⁺ are Y₁ = 6.1004X + 323.5808; Y₂ = 0.5861X + 2468.1944 and Y₃

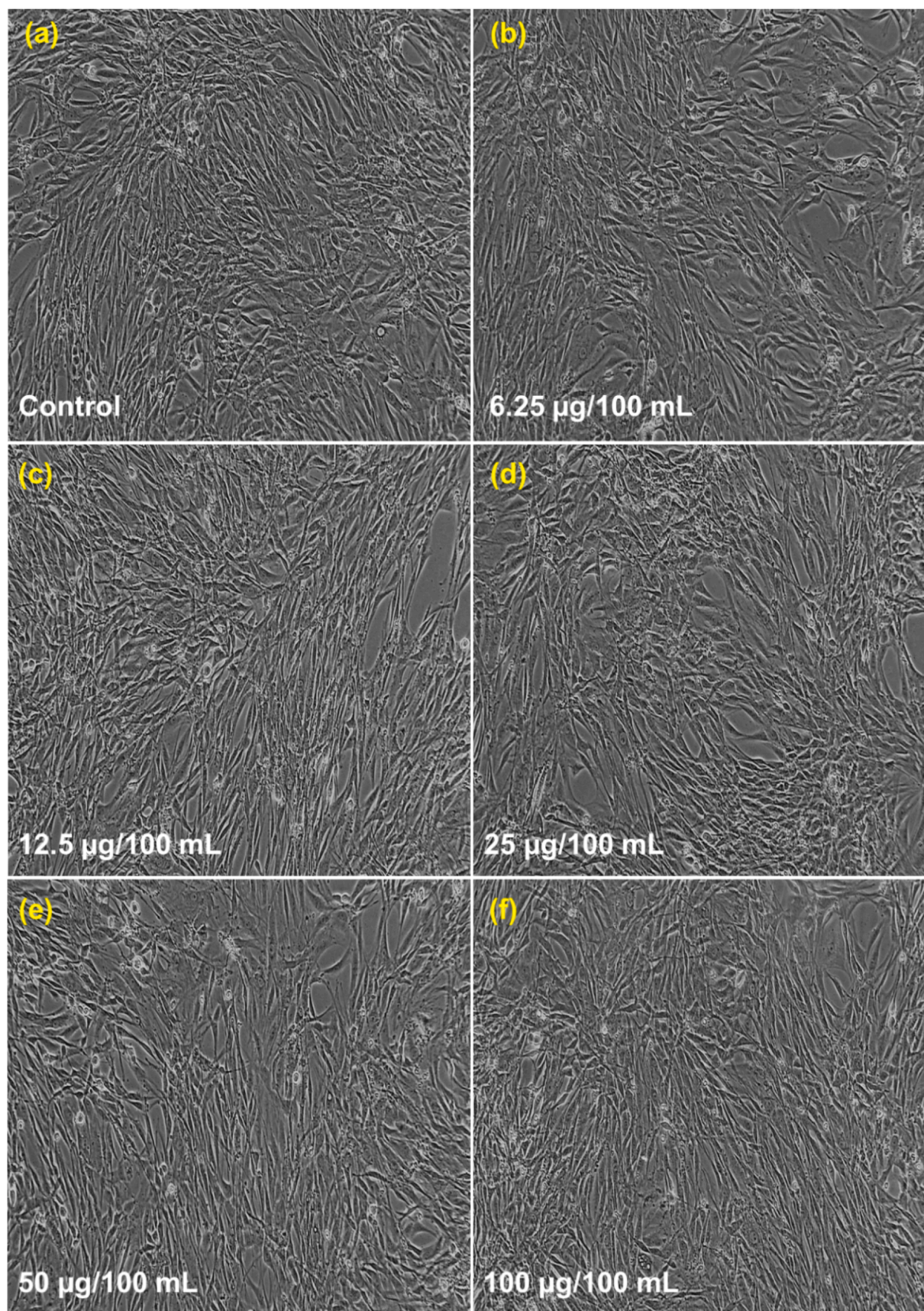


Fig. 16. Cytotoxic results for CeNC@T1: (a) Control, (b)-(f) Different concentrations of CeNC@T1 (6.25–100 µg/mL) incubated with L6 cell line.

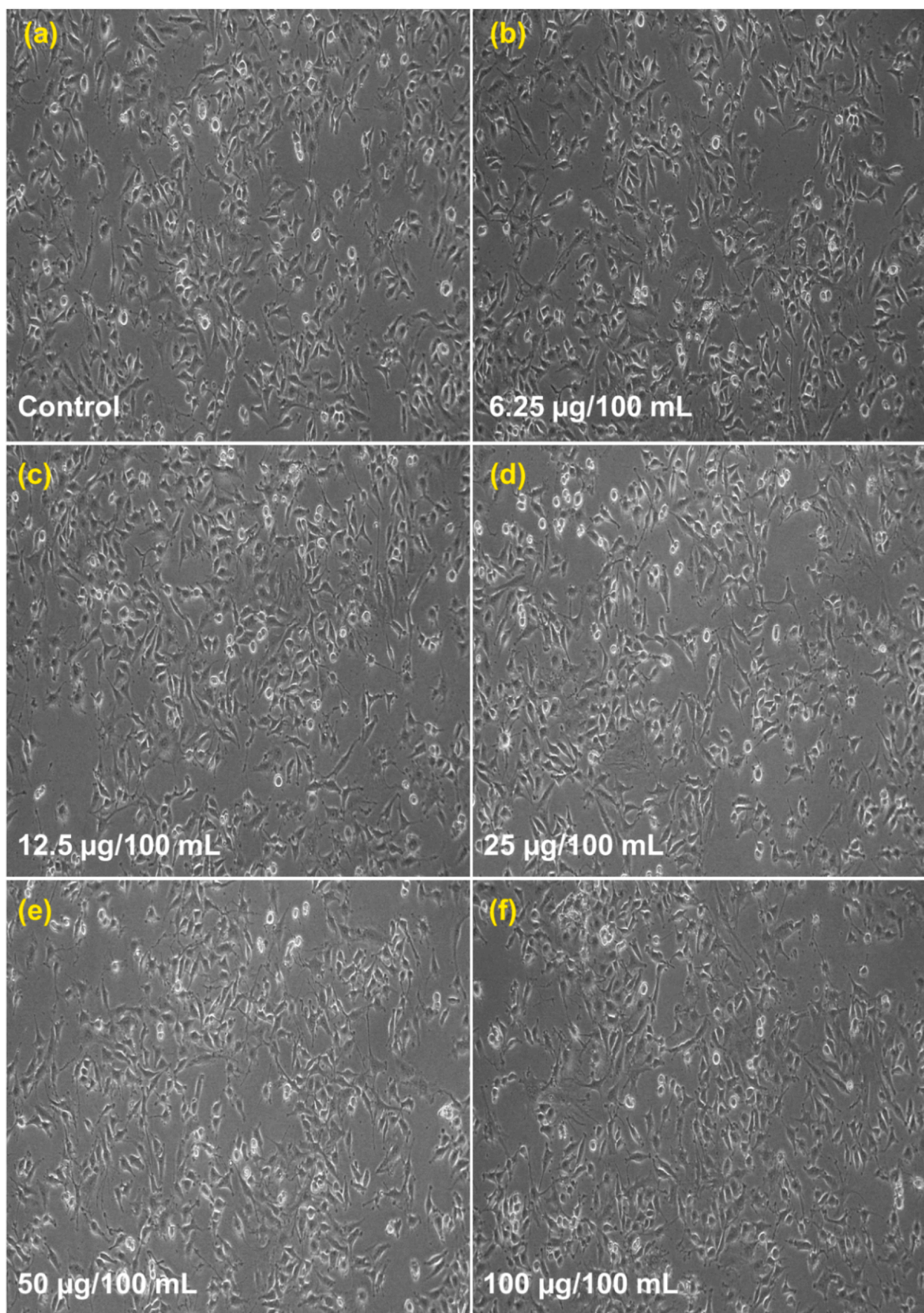


Fig. 17. Cytotoxic results for CeNC@T2: (a) Control, (b)-(f) Different concentrations of CeNC@T2 (6.25–100 µg/mL) incubated with L6 cell line.

= 0.06173X + 5091.25 with a correlation coefficient of 0.9890, 0.9993 and 0.9915 respectively, where X is the concentration of Cd²⁺ (nM). The LOD is calculated to be 2.56x10⁻⁸ M, 2.66x10⁻⁷ M, 2.53x10⁻⁶ M. The regression equations for Hg²⁺ are: Y₁ = 5.02132X + 324.3137; Y₂ = 0.4211X + 1932.3333; Y₃ = 0.06484X + 3889.4643 with a correlation coefficient of 0.9834, 0.9946 and 0.9917 respectively, where X is the concentration of Hg²⁺ (nM). The LOD is 3.11x10⁻⁸, 3.71x10⁻⁷, 2.41 x10⁻⁶ M.

The fluorescence intensity of Ag-CeNC@T3 is also greatly enhanced in the presence of Hg²⁺, Pb²⁺ and Cd²⁺ and the increase is much higher than CeNC@T3 (Fig. 13). On mixing the Hg²⁺ ions to an aqueous dispersion of Ag-CeNC@T3, the emission intensity is directly proportional to [Pb²⁺]. Three linear regions are observed on varying the concentration of Hg²⁺ from 5.0x10⁻⁹ M to 4.0x10⁻⁵ M. The corresponding regression equations are Y₁ = 9.4062X + 1363.4360; Y₂ = 0.8473X + 4118.3333; Y₃ = 0.1366X + 7664.2857 with a correlation coefficient of 0.9823, 0.9883 and 0.9953 respectively, where X represents the concentration of Hg²⁺ (nM). The LOD is 10.13x10⁻⁸ M, 11.24x10⁻⁷ M, 6.97 x10⁻⁶ M.

Pb²⁺ and Cd²⁺ ions also increase the emission intensity to a greater extent. In these cases also, there are three linear regions in which the emission intensity is directly proportional while increasing the concentration of these ions 10,000 folds between 5.0x10⁻⁹ M to 4.0x10⁻⁵ M. The regression equations for Pb²⁺ are Y₁ = 8.9245X + 1293.9777; Y₂ = 0.7873X + 4206.1111; Y₃ = 0.1091X + 7683.9286 with a correlation coefficient of 0.9750, 0.9979 and 0.9990 respectively, where X is the concentration of Pb²⁺ (nM). The Limit of Detection in the three linear regions is found to be 10.68x10⁻⁸, 1.21x10⁻⁶, 8.73x10⁻⁶ M. The regression equations for Cd²⁺ are Y₁ = 9.02267X + 1185.5860; Y₂ = 0.4893X + 4543.333; Y₃ = 0.05636X + 6455.7143 with a correlation coefficient of 0.9819, 0.9975 and 0.9807 respectively, where X is the concentration of Cd²⁺ (nM). The LOD is estimated as 10.56x10⁻⁸ M, 1.95x10⁻⁶ M, 1.69x10⁻⁵ M in the three different regions.

Various mechanisms have been reported in literature for the sensing of heavy metals by CePO₄ based nanocomposites. The coordination of Pb²⁺ ions with phosphate group forming an insoluble complex is reported to be the reason for the quenching of fluorescence of cerium phosphate (Dezfuli et al., 2014); (Khandare et al., 2014); (Song et al., 2019). The presence of Hg(II) leads to enhanced fluorescence in bio-functionalized Ag nanoparticles. This is a result of fluorescence resonance energy transfer, facilitated by the aggregation of nanoparticles brought about by the complexation of Hg(II) with Ag nanoparticles (Sebastian et al., 2018). A similar pattern has been documented in the instance of Ru-loaded CeO₂ NCs for electrochemical detection of Hg(II), attributed to an abundance of oxygen vacancies (Sun et al., 2020). In this case, it appears that the heavy metals are adsorbed over the oxygen vacancies created in the NCs, enhancing fluorescence. As the availability of vacancies decreases, the intensity of fluorescence also decreases, leading to different linear regions.

4.1.1. Turn-on-off sensor

In order to see the efficiency of the NC as "on - off" sensors, emission studies were carried out first with fixed concentration of lead ions ([Pb²⁺] = 5.0x10⁻⁶ M) and then mixing different concentrations of EDTA by keeping the total volume as constant. The intensity of emission decreased with increasing [EDTA] up to 99.88 %. The results are represented in Fig. S8. This indicates that EDTA quenches the emission of CeNC@T3 which is enhanced in presence of Pb²⁺ ions. To revive the fluorescence, emission studies were again carried out with solutions which were quenched with three different concentrations of EDTA in turn-off experiments by adding 5.0x10⁻⁴ M concentration of Pb²⁺ ions. With the addition of Pb²⁺, the emission is again revived. The results are produced in Fig. S9. To study the suitability of CeNC@T3 to sense real-world sample, an aqueous solution which is a mixture of three heavy metals viz., Pb²⁺, Cd²⁺ and Hg²⁺ was mixed with CeNC@T3 and the emission was measured. The emission intensity has increased but not as

additive as that of all ions. The results are consolidated in Fig. S10.

4.2. Evaluation of SPF of CeNC@T1 and CeNC@T2

Commonly, sunscreens are protectants that either absorb or block the harmful UV rays from sunlight and show immunosuppressive effects. In summer, using sunscreen is an efficient way for skin protection, which lessens UV- induced ROS-mediated photoaging. SPF is defined as the ratio of the required UV energy to generate a Minimum Erythema Dose (MED) on protected skin to the UV energy essential to produce MED that is not protected. MED is the shortest span of time or amount of UV light exposure required to cause the most perceptible and minimal erythema on exposed skin. SPF number determines the efficacy of a sunscreen-based product. Sunscreens are more effective at preventing sunburns if their SPF value is higher.

The results of SPF at different concentrations of CeNC@T1 are given in Fig. 14. SPF values exhibit a marked increase with increasing concentration of NC. The possible mechanism for the good UV shielding property of NC can be due to their smaller size (2.36 nm), larger band gap (E_g = 3.03 eV) and photon energy consumption. The SPF of CeNC@T2 is also given in the figure for comparison. CeNC@T2 demonstrates less SPF in comparison to CeNC@T1, this may be due to their larger size and less band gap (E_g = 2.95 eV) compared to CeNC@T1 and photon energy consumption (Miri et al., 2020). High SPF values are observed even with low concentrations (100 µg/mL), making this system more significant than those reported in literature (Table 1).

4.3. Evaluation of the cytotoxic effect of CeNC@T1 and CeNC@T2

The variation of percentage cell viability with different concentrations of CeNC@T1 and CeNC@T2 are shown in Fig. 15. It is found that a concentration of 6.25–25 µg/mL yields a viability rate of 99 %. At 50 and 100 µg/mL concentration, it is 98.62 % and 97.93 % for CeNC@T1 and 98.15 % and 97.70 % for CeNC@T2 respectively, as given in Figs. 16 & 17. According to the findings of cytotoxicity studies, L6 cell line is not significantly harmed by the marine shrimp produced CeNC@T1 and CeNC@T2 and shows no notable toxicity (Miri et al., 2021).

5. Conclusion

This article introduces an uncomplicated, innovative and eco-friendly method for synthesizing pure CeO₂-CePO₄ NCs and silver-doped CeO₂-CePO₄ NCs, utilizing the edible marine shrimp *P. semisulcatus*. This process paves the way for a fresh approach in the production of various types of bio-nanocomposites. CeO₂ exhibits properties of a wide-band semiconductor and owing to the presence of Ce³⁺ ions stemming from oxygen vacancies, both direct and indirect band-to-band transitions can be observed in CeO₂ nanoparticles. The direct bandgap of CePO₄ nanoparticles measures approximately 3.0 eV. Thanks to these band gaps, these nanocomposites function as effective UV shields, and their low toxicity makes them suitable for sunscreen applications. Furthermore, their extended excited state lifetimes enable their efficient use as fluorescence sensors for heavy metal ions. The broad range of applications for these synthesized nanocomposites positions them as highly promising materials for future.

Declaration of Competing Interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

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Appendix A. Supplementary data

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