

21 ST INTERNATIONAL CONFERENCE ON FUSION REACTOR MATERIALS ICFRM-21 GRANADA, SPAIN 22ND – 27TH OCTOBER, 2023



Optimisation of ECX permeation barriers towards thicker alumina

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This work has been carried out within the framework of the EUROfusion Consortium, funded by the European Union via the Euratom Research and Training Programme (Grant Agreement No 101052200 — EUROfusion). Views and opinions expressed are however those of the author(s) only and do not necessarily reflect those of the European Union or the European Commission. Neither the European Union nor the European Commission can be held responsible for them.

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NO TRITIUM, NO FUSION

Tritium to maintain the fusion reaction best produced right in the fusion reactor (breeding blanket).

Small isotope that penetrates or permeates through structural materials, especially steel.

Hazardous substance that should arrive at the reactor unit for controlled tritium extraction.

In the breeding blanket, tritium-permeation barrier (TPB) on structural materials required, on the side *Harvesting energy fro deuterium and tritium.* of the tritium source.



Lithium

Breeding

blanket

Neutrons

🗿 🚳 🕋

Deuterium

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Electricity or



TPB CANDIDATE MATERIALS

Low hydrogen solubility and diffusivity, paired to high hydrogen retention.

- In the breeding blanket:
- Tolerant of unprecedented neutron irradiation, minimum of neutron absorption.
- Chemically resistant against the breeding material/atmosphere.
- Thermo-mechanically compatible with the substrate material, e.g., thermal cycling.

Appropriate deposition process for producing a defect-free coating must exist.



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Oxides, Nitrides, Carbides, Some metals, e.g., tungsten.



ALUMINA

Favourable properties as to hydrogen permeation, especially for α -Al₂O₃.

Issues as to service in the breeding blanket:

- Anisotropic swelling of hexagonal α -Al₂O₃, of less concern for cubic γ -Al₂O₃ or amorphous Al₂O₃.
- Aluminium tends to be activated, needs to be removed from the components after service.
- Lithium uptake from the breeding material, especially in the case of eutectic lithium–lead.

Unsolved issues relating to neutron irradiation irrelevant to the tritium-extraction and removal sub-system (TER).







Ball-and-stick model of α–Al₂O₃. By Ben Mills - Own work, Public Domain, https://commons.wikimedia.org/w/index.php?curid=4222743



ECX PROCESS: (1) ELECTROCHEMICAL X (AL)





Aluminium transfer from a dissolving anode sheet onto the (EUROFER) substrate.

DEPOSITION

Ionic liquid used as an electrolyte (1-Ethyl-3-methyl-imidazolium-chloride + AICl₃,1:1.5).

15 mA per cm² of substrate surface area (deposition rate ~20 μ m Al/h), at 110 °C, in dry argon atmosphere (< 1 ppm O₂ or H₂O).

Pulse plating with 1 Hz.



Set-up for electroplating.

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Heat treatment below aluminium melting point in technical argon (640 °C, 4 h)

InterdiffusionAluminides, brittleOxides nucleate



High-temperature heat treatment in oxidising atmosphere (980 °C, 0.5 h)

Transformation of aluminidesNucleation and growth of alumina



To produce aluminium-diffusion coating with alumina surface film.

Standard oxidising atmosphere in stage 2 and 3: Technical argon.

Heat treatment to produce an aluminium-diffusion coating with alumina surface film.





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RAISING THICKER ALUMINA: WHY AND HOW?



Standard heat treatment in argon:

- Corrosion resistance in flowing eutectic lithium– lead suggests thin alumina surface layer, but
- Too thin for satisfactory (deuterium) permeation reduction.





Surface recession (µm) after exposure to flowing Pb–16Li at 550 °C (PICOLO) of ECX coated EUROFER in contrast to RAFM steels and cross section of ECX–coated EUROFER after exposure for 10,000 h.

(S.-E- Wulf et al., Workshop on LM technology, Nov. 6th-7th 2018, Madrid, Spain).

Approach to thicker alumina:

- Change heat-treatment atmosphere during stage 2 and 3 to air.
- Prolong holding at 980 °C.

	640 °C	980 °C	760 °C
Standard	4 h, Ar, C1	0.5 h, Ar, C1	1.5 h, Ar, C1
Mod. 1	4 h, Ar, C1	0.5 h, air , C2	1.5 h, air , C2
Mod. 2	4 h, Ar, C1	<u>1 h</u> , air , C2	1.5 h, air , C2
Mod. 3	4 h, Ar, C1	<u>2 h</u> , air , C2	1.5 h, air , C2

Time, atmosphere, cooling (C1: in cold furnace section, Ar stream; C2: air cooling outside furnace).



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SAMPLES AND CHARACTERISATION



Standard heat-treatment in argon and three modified heat-treatments.

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- Ø35×1 mm discs for later tests on deuterium permeation (at CIEMAT), X-ray diffraction spectroscopy (XRD) and metallographic examination.
- Ø8×24 mm cylinders for exposure to eutectic lithium– lead in the PICOLO loop (550 °C, low flow velocity to static).

+ Additional flat sample for XRD before and after each stage of the standard heat treatment





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Ø8×24 mm cylinder for tests in PICOLO loop.



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METALLOGRAPHIC ANALYSIS OF THE COATINGS

Scanning electron microscopy (SEM) on cross sections, at medium magnification, combined with energy-dispersive X-ray spectroscopy (EDS)

- ✤ 1–2 µm oxide for second and third stage of heat treatment in air, $<< 1 \mu m$ for standard treatment in argon.
- Oxide clearly aluminium-rich. •
- Oxide protrusions interlock the surface layer with the * underlying metal, with benefit to oxide adherence.
- 40 and 70 μm aluminium penetration into the substrate steel after 0.5 and 2 h holding at 980 °C, respectively.





(c) 1 h at 980 °C, air





Surface oxide on ECX-coated EUROFER with (a) standard heat treatment in argon and (b-d) modified heat treatment.











- Phase transitions in the metallic portion of the ECX coating.
- After standard heat treatment, any oxide possibly formed too thin for assessment with XRD.

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- Oxide signals appear after modified heat treatments, especially α–Al₂O₃.
- Intensity of peaks associated to oxides increases with time of holding at 980 °C.



EXPOSURE TO PB-16LI AT 550 °C

After 3000 h in the PICOLO loop, under lowflow to stagnant conditions

- Metallic part of the coating largely protected from contact with the liquid metal.
- Number of weak spots the smaller the longer the time of holding at 980 °C in air.
- Oxide layer heterogeneous, locally detached from the substrate (?).
- Around 50 µm aluminium penetration into the substrate irrespective of the heat treatment.





Electron-optical micrographs of ECX coating with heat treatment Mod. 2 (Table 1) after exposure for 3000 h to eutectic Pb–16Li.



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CONCLUSIONS



- Anticipated effect of air as to promoting alumina growth, but optimisation for continuity or minimisation of defects still required.
- Prolong holding at 980 °C? Probably not an option for EUROFER substrate.
- Especially for ECX, other selective changes with regard to the heat-treatment rather than the deposited aluminium (e.g., doping with other elements).
- Why not grow γ -Al₂O₃ then?





ACKNOWLEDGEMENTS



The authors would like to thank their colleagues Drs. A. Heinzel (KIT-IHM) and K. Seemann (KIT-IAM-AWP) for conducting and evaluating XRD measurements.

Financial support by the Nuclear Fusion Programme (FUSION) of KIT is gratefully acknowledged.





This work has been carried out within the framework of the EUROfusion Consortium, funded by the European Union via the Euratom Research and Training Programme (Grant Agreement No 101052200 — EUROfusion). Views and opinions expressed are however those of the author(s) only and do not necessarily reflect those of the European Union or the European Commission. Neither the European Union nor the European Commission can be held responsible for them.



