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## Article

# Large-Scale Ex Situ Tests for CO<sub>2</sub> Storage in Coal Beds

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**Abstract:** This publication discusses the experiments and findings of project ROCCS (Establishing a Research Observatory to Unlock European Coal Seams for Carbon Dioxide Storage), which aimed to investigate the potential for carbon dioxide storage in coal seams. The project involved large-scale ex situ laboratory tests, where CO<sub>2</sub> was injected into an experimental coal seam using a high-pressure reactor at the Central Mining Institute in Poland. The reactor simulated underground conditions, and the experimental coal seam measured 3.05 m in length with a cross-section of 0.4 × 0.4 m. Parameters such as gas flow, temperatures, and pressures were monitored during the experiments. In the study conducted, the sorption capacity of coal from the Polish mine “Piast-Ziemowit” for CO<sub>2</sub>, at a sorption pressure of 30 bar, was determined to be 4.8% by weight relative to the raw coal mass. The data collected from these ex situ tests can support the design of a potential commercial-scale CO<sub>2</sub> storage installation.

**Keywords:** CO<sub>2</sub> storage; coal; horizontal well; ex situ test; adsorption



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## 1. Introduction

CO<sub>2</sub> is considered one of the main causes of the occurrence of global climate change [1]. Adverse environmental changes such as rising sea levels, melting glaciers droughts, and hurricanes can be observed in many parts of the world [2]. Over the past few decades, an increase in the amount of carbon dioxide in the air has contributed to rising global temperatures [3]. Many countries are leaning into the problem of global warming and trying to find ways to slow down the process [4,5]. CO<sub>2</sub> is emitted mainly from the combustion of fossil fuels, both in large combustion units, such as those used to generate electricity, and in smaller, distributed sources, such as car engines and furnaces used in residential and commercial buildings.

To improve the climate situation, it is necessary to reduce the amount of CO<sub>2</sub> in the atmosphere [6,7]. It is possible to reduce levels. However, this involves changing energy technologies, which is an expensive undertaking.

One way to solve the problem of excess CO<sub>2</sub> is to collect and locate it in geological structures. In the study [8], the possibilities of placing CO<sub>2</sub> in oil or gas deposits and saline aquifers were shown. The high potential for CO<sub>2</sub> storage in coal seams is also known—coal’s inherent adsorption properties, derived from its porous structure and abundant micropores, enable it to effectively bind and store various gases on its internal surfaces [9–11]. Carbon Dioxide-Enhanced Coal Bed Methane (CO<sub>2</sub>-ECBM) is a technology where CO<sub>2</sub> is injected into coal seams to displace and enhance the production of methane while simultaneously storing the CO<sub>2</sub> underground. The process leverages coal’s higher affinity for CO<sub>2</sub> over methane, facilitating both methane recovery and carbon sequestration [12–15]. In their study, a team led by Zheng showed a designed large-scale device to

simulate CO<sub>2</sub>-ECBM technology. The results of the study were that dynamic injection of CO<sub>2</sub> into the reservoir results in more efficient CH<sub>4</sub> recovery and CO<sub>2</sub> sequestration [16,17]. Similar studies were conducted by the Bai team [18]. The research was aimed at learning the relationship between CO<sub>2</sub> injection and CH<sub>4</sub> desorption. It was observed that, in the process of CH<sub>4</sub> displacement by CO<sub>2</sub>, the temperature of the coal seam increases. In the study conducted by Wang [4], investigations on changes in the coal structure during CO<sub>2</sub> injection tests were conducted. The effect of CO<sub>2</sub> intrusion on the properties of coal as a result of changes in molecular structure and CO<sub>2</sub> capture mechanisms was also investigated.

Another attempt to explain the complex CO<sub>2</sub> sorption process was made by Harinandan Kumar [19]. His study attempted to investigate the sorption capacity of Indian coals. The result was the development of a sorption isotherm. The study showed that sorption capacity increased with increasing pressure, but decreased at higher temperatures.

In the study [20], the impact of underground CO<sub>2</sub> storage on a coal deposit was investigated using computer simulations. The research focused on hydro-mechanical interactions concerning Polish coals located in the Upper Silesian Energy Region. Computer simulations showed that, under certain conditions, the CO<sub>2</sub> injection process is safe and has no negative environmental effects. However, when the CO<sub>2</sub> injection pressure exceeded 1.5 times the hydrostatic pressure, environmental hazards such as fault reactivation could emerge. The consequence of this phenomenon could lead to damage to surface infrastructure.

In the paper [21], it was found that, in the case of attempts to store CO<sub>2</sub> in a coal bed, the problem of coal swelling arises. This is an unfavorable phenomenon that leads to the formation of a zone of high CO<sub>2</sub> concentration near the injection well.

In the project with the acronym CO<sub>2</sub>RINA (CO<sub>2</sub> Storage Risk Integrated Analysis), risks associated with the storage of CO<sub>2</sub> in underground coal deposits were analyzed. With this approach, it is possible to study in more detail the complex processes occurring in deposits when storing CO<sub>2</sub>. More information on this subject can be found in reference [22].

The ROCCS project builds upon prior research, including the RECO<sub>2</sub>POL project conducted by the Central Mining Institute and funded by the European Commission [23,24]. The experiment, which lasted several months, involved the injection of 760 tons of CO<sub>2</sub> into a coal seam using a vertical injection well. The main objective of the research was to determine whether the injected CO<sub>2</sub> was adsorbed onto the coal or remained as a free gas. The results presented indicate that CO<sub>2</sub> was adsorbed onto the coal—however, it was found that the required CO<sub>2</sub> injection pressures were higher than anticipated, with CO<sub>2</sub> injection only possible through cycles of pressure build-up and fall-off. The permeability of the coal seam reduced with time due to CO<sub>2</sub> sorption-induced coal swelling, and it was only possible to maintain continuous injection after hydraulic fracturing was completed. ROCCS acts on the recommendation of RECO<sub>2</sub>POL that “other well completions, such as horizontal or ‘fishbone’ drilling, need to be researched to assess their impact on injectivity and productivity”.

Several projects have highlighted a significant theoretical potential of CO<sub>2</sub> storage and usage for coal bed methane (CBM) recovery [25,26]. They have also indicated that this potential is constrained by the fact that European coal seams are significantly less permeable than those in the USA, where the largest field trials have taken place. Combined with permeability losses due to CO<sub>2</sub> sorption-induced coal swelling, a principal concern is therefore the need to over-pressurize coal seams to sustain injection, leading to concerns over caprock integrity and leakage. ROCCS acts on these constraints by determining the potential of horizontal injection wells to deliver sustained injection rates, circumventing the need for significant over-pressurization.

The reliability of experimental assessment and practical application scenarios of in situ experiments to assess the adsorption capacity of coal depends on various factors such as the quality of the data, the accuracy of the measurements, and the validity of the assumptions made during the experiments. The results of these experiments can be used to estimate the amount of carbon dioxide that can be stored in coal seams, but they are subject to uncertainty and variability. Some studies have shown that the adsorption capacity of coal

can vary depending on factors such as temperature, pressure, and coal type [27,28]. Other studies have suggested that the adsorption capacity of coal can be affected by factors such as SO<sub>2</sub> and pH concentration [29]. Overall, while in situ experiments can provide valuable information about the adsorption capacity of coal, they should be interpreted with caution and used in conjunction with other methods to estimate carbon storage potential.

To complement the in situ experiment, which exhibited the aforementioned limitations, a controlled laboratory test was conducted. The high-pressure laboratory installation enabled the simulation of CO<sub>2</sub> injection in underground conditions at the surface (ex situ) into an experimental coal seam. This setup aimed to determine the coal's sorption capacity and observe the dynamics of CO<sub>2</sub> migration. The large coal blocks were placed inside the insulated reactor, preventing gas losses outside the experimental zone. The size of the experimental coal seam, combined with the reactor's capability to operate at a high pressure of 30 bar, allowed for the simulation of coal–gas interactions. A longitudinal sorption channel was cut in the lower part of the prepared coal seam to emulate the design of a horizontal injection well. Inlet and outlet gas flow, temperatures, and pressures were also monitored as crucial operational parameters.

The results of the ex situ CO<sub>2</sub> injection test, combined with other laboratory investigations and numerical modeling, offer valuable insights. Together with in situ tests, they can guide the design of a commercial-scale installation. This installation could potentially be located within one of the Polish mines for large-scale CO<sub>2</sub> storage purposes.

## 2. Materials and Methods

### 2.1. Coal Properties and Sample Preparation

The material used for sorption testing was hard coal from the “Piast-Ziemowit” mine. The results of proximate and ultimate analyses for the coal sample under study are presented in Table 1. All analyses were carried out by the Department of Solid Fuels Quality Assessment of the Central Mining Institute (PN-EN ISO/IEC 17025:2018-2 accreditation certificate [30]) in Katowice. The hard coal used is characterized by a moderate moisture and ash content, approximately 8.50% and 7.56%, respectively, (as received).

**Table 1.** Proximate and ultimate characteristics of “Piast-Ziemowit” coal (unit is % by weight if not printed otherwise).

	Parameter	Symbol	Unit	Value	Standard <sup>b</sup>
As received	Total moisture	W <sub>t</sub> <sup>r</sup>	(%)	8.50	PN-G-04560:1998 [31]
	Ash	A <sup>r</sup>	(%)	7.56	PN-G-04560:1998
	Volatile matter	V <sup>r</sup>	(%)	30.15	PN-G-04516:1998 [32]
	S total	S <sub>t</sub> <sup>r</sup>	(%)	0.98	PN-G-04584:2001 [33]
	Lower heating value	Q <sup>r</sup>	(kJ/kg)	25,786	PN-G-04513:1981 [34]
Analytical	Moisture	W <sup>a</sup>	(%)	7.47	PN-G-04560:1998
	Ash	A <sup>a</sup>	(%)	7.64	PN-G-04560:1998
	Volatile matter	V <sup>a</sup>	(%)	30.49	PN-G-04516:1998
	Lower heating value	Q <sup>a</sup>	(kJ/kg)	26,103	PN-G-04513:1981
	S <sub>t</sub> <sup>a</sup> total		(%)	0.99	PN-G-04584:2001
	C <sup>a</sup>		(%)	68.62	PN-G-04571:1998 [35]
	H <sup>a</sup>		(%)	4.30	PN-G-04571:1998
	N <sup>a</sup>		(%)	1.08	PN-G-04571:1998
	O <sup>a</sup> *		(%)	10.20	

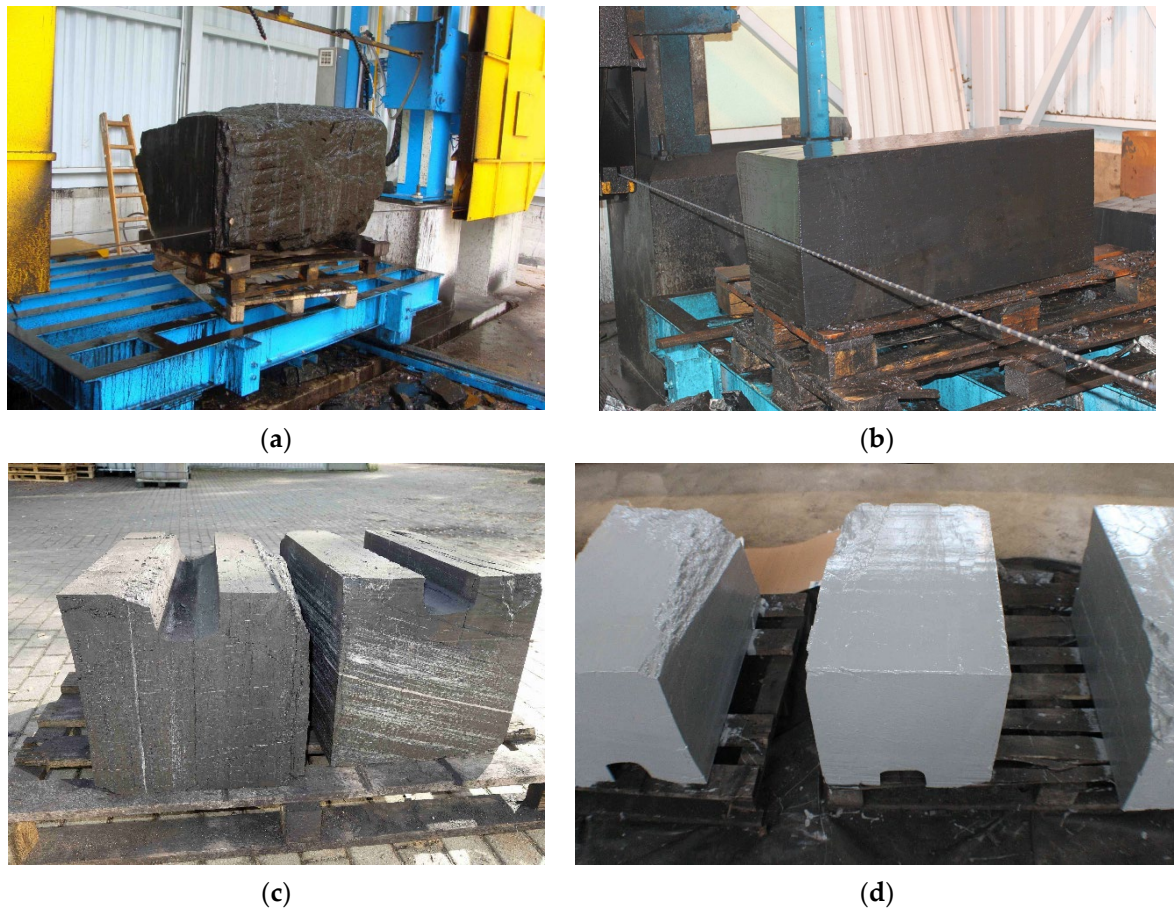
<sup>r</sup>—as received, <sub>t</sub>—total, <sup>a</sup>—analytical, and <sup>b</sup>—Polish testing by the accredited laboratory. \* Oxygen calculated as: (O<sup>a</sup>) = 100 − (W<sup>a</sup>) − (A<sup>a</sup>) − (C<sup>a</sup>) − (H<sup>a</sup>) − (S<sub>t</sub><sup>a</sup>) − (N<sup>a</sup>). S total—sum of inorganic and organic Sulphur content.

Before the coal was used as a sorbent in a pressurized vessel (a reactor), the acquired coal blocks had to be properly cut. Longitudinal sorption channels were cut in the lower part of the prepared coal blocks to allow the gas to be sorbed to penetrate. The cross-section of the cut coal blocks was 0.4 × 0.4 m, and the sorption channel was approximately 0.15 × 0.15 m. Five blocks of coal were cut to lengths of 0.31, 0.74, 0.70, 0.60, and 0.70 m,



respectively, based on the initial size and shape of the acquired coal samples. These blocks were combined to form an artificial coal seam with a total length of 3.05 m. The weight of all the blocks was 631 kg.

The procedure of cutting coal blocks is presented in Figure 1, while the arrangement and shapes of the used coal blocks are shown in Figure 2 (the proportions of sides were maintained).



**Figure 1.** Preparation of artificial coal seam for sorption tests: (a,b) sample cutting, (c) cut blocks of coal with a sorption channel, (d) fully painted blocks of coal prepared for sorption testing.



**Figure 2.** Characteristics of an artificial coal seam for sorption testing. The lengths of coal blocks 1 to 5 are 0.31 m, 0.74 m, 0.70 m, 0.60 m, and 0.70 m, respectively.

The dimensions, weights of the coal blocks, and their volume are given in Table 2. The volumes of coal blocks were calculated after subtracting the volume of the excised sorption channel. The total volume of the coal blocks inserted into the reactor was 458 dm<sup>3</sup>.

**Table 2.** Dimensions and weights of coal blocks used.

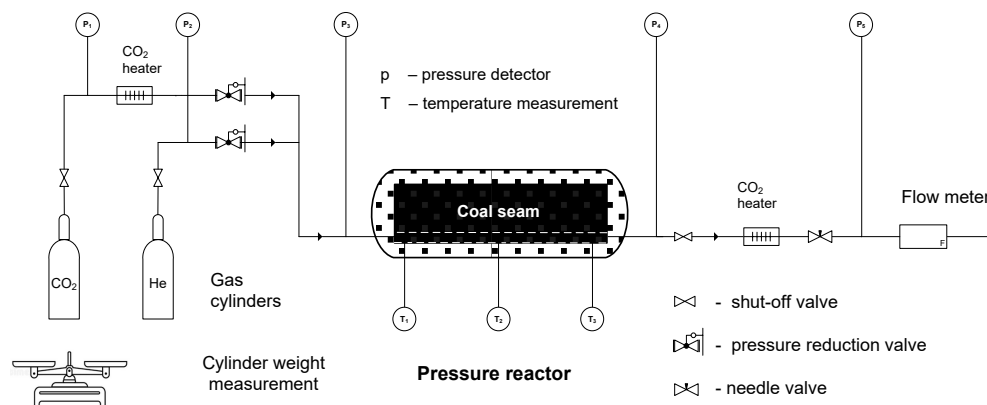
Coal Block	Width (cm)	Height (cm)	Length (cm)	Weight (kg)	Volume of a Coal Block (dm <sup>3</sup> )
1	40	40	31	58	47.4
2	40	40	74	139	108.0
3	40	40	70	184	101.3
4	40	40	60	106	96.2
5	40	40	70	144	105.1
TOTAL				631	458.0

Following this, with the exception of the sorption channels, all surfaces of each coal block were painted several times (Figure 1d) with a gas- and moisture-impermeable varnish. The objective was to ensure that the sorption process occurred exclusively through the sorption channel. The type of coating used was selected based on preliminary studies. For each test, small blocks of coal were painted with different types of lacquer. After drying, their permeability was assessed by immersing the painted blocks in water for 24 h. The best isolation was achieved using a chloride–rubber varnish produced by a Polish company. This varnish is commonly used to protect car chassis from corrosion.

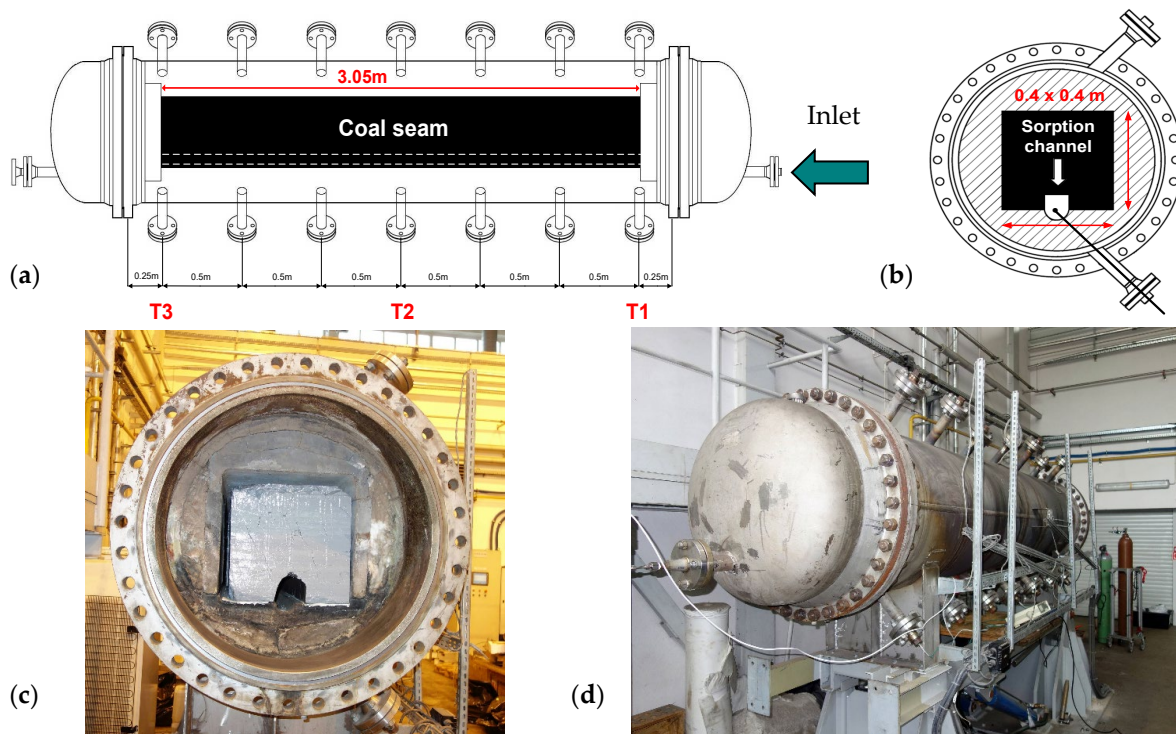
2.2. Experimental Setup

A scheme of the installation used for sorption testing is shown in Figure 3, while cross-sections and a view of the reactor are shown in Figure 4.

The measuring elements of the thermocouples were placed so that they measured temperatures in the center of the sorption channel. The measurement accuracy of analog pressure sensors P1–P3 was ±1 bar, while that of sensor P5 was ±0.1 bar. The measurement accuracy of digital pressure sensor P4 was 0.01 bar, while that of digital temperature sensors T1–T3 was 0.01 °C, respectively. Digital data values were recorded automatically every 20 s. A bellows gas meter from ELSTER, type BK-G10M, was used to measure the amount of gas flowing out of the reactor. The P5 pressure sensor was used to control the expansion pressure released from the helium or carbon dioxide reactor. The scale seen in Figure 3 was used to control the weight of the gas cylinders before and after sorption. The difference in the weight of the cylinders gave information about the amount of gas released from the cylinder. The accuracy of the cylinder weight measurement was 0.1 kg. The expansion of the compressed CO<sub>2</sub> causes a temperature drop—as a result, valves placed in the path of the gaseous CO<sub>2</sub> might have frozen. The use of dedicated heaters prevented this adverse effect.



**Figure 3.** Scheme of the ex situ pressure sorption installation.



**Figure 4.** Pressure reactor for sorption testing: (a) side cross-section, (b) vertical cross-section, (c) reactor with coal seam inside, (d) reactor ready for the CO<sub>2</sub> sorption test.

### 2.3. Experimental Procedure

The experimental procedure consisted of two stages. Stage I involved determining the free volume in the reactor using helium introduced into the reactor at selected pressures of 3 and 6 bar. In Stage II, the sorption capacity of raw coal for absorbing carbon dioxide was measured under a pressure of 30 bar.

The following activities and measurements were required for Stage I and II:

Stage I—helium tests:

- Installation of the measurement system and helium filling system (from gas cylinder) to the reactor,
- Measurement of the temperature and pressure prevailing in the gas cylinder before filling the reactor with helium,
- Filling the reactor with helium to a pressure of 3 bar and waiting until the gas pressure in the reactor is established,
- Depressurization of the reactor to atmospheric by opening the outlet and needle valves and releasing helium into the atmosphere,
- Measuring the volume of helium released from the reactor with a flow meter,
- Re-measuring the temperature and pressure of the remaining helium in the gas cylinder,
- Repeating the procedure for a higher pressure of 6 bar.

Stage II—carbon dioxide test:

- Connection of the measurement system and CO<sub>2</sub> filling system (gas cylinder) to the reactor,
- Measuring the mass of a liquid CO<sub>2</sub> cylinder before filling the reactor with carbon dioxide,
- Filling the reactor with CO<sub>2</sub> to a starting pressure of 30 bar. Close the gas cylinder valve and wait until the gas pressure in the pressurized reactor is established,
- Calculating, from the difference in the weight of the cylinder, the volume of CO<sub>2</sub> let into the reactor at 30 bar. Reducing the pressure in the reactor to atmospheric by opening the outlet needle valve and releasing the CO<sub>2</sub> into the atmosphere,
- Measuring the volume of CO<sub>2</sub> released into the atmosphere with a flow meter,

- Calculating the difference in the mass of CO<sub>2</sub> contained in a full and used gas cylinder and converting the mass to gas volume. The difference in CO<sub>2</sub> volume calculated in is the amount of CO<sub>2</sub> adsorbed.

### 3. Results

#### 3.1. Helium Tests

##### 3.1.1. Measurement of the Free Volume of the Reactor without the Coal Block Located Inside

The equation of state of a perfect gas (1) and Clapeyron's Equation (2) were used to determine the free volume in the reactor:

$$\frac{pV}{T} = \text{const} \quad (1)$$

$$pV = nRT \quad (2)$$

where: p—pressure, V—volume, T—temperature, R—universal gas constant, n—number of moles.

First, the volume of helium taken from the gas cylinders during the filling of the pressurized reactor was calculated. Two measurement tests were performed at 3 and 5 bar. Each test was performed using a separate helium gas cylinder. Due to insufficient gas contained in the helium cylinder during the second test, the planned reactor pressure of 6 bar needed to be lowered to 5 bar. The results of the measurements and calculations converted to standard conditions ( $T = 273.15 \text{ K}$ ,  $p = 1013 \text{ hPa}$ ) are shown in Table 3.

**Table 3.** Calculation of helium volume taken from gas cylinders.

Test No.	Initial Pressure in the Helium Cylinder	Final Pressure in the Helium Cylinder	Helium Cylinder Volume	Average Helium Cylinder Temperature	Amount of Helium Taken from the Gas Cylinder	
	bar	bar		°C	mol	Ndm <sup>3</sup>
1 (3 bar)	198	89	50	5.5	235.23	5269.3
2 (5 bar)	181	5.0	50	2.9	382.16	8560.4

From the data presented in Table 3, the helium volume taken for the test at pressures of 3 and 5 bar was 5269 Ndm<sup>3</sup> and 8560 Ndm<sup>3</sup>, respectively (under standard conditions). The gas taken from the helium cylinder was introduced into a pressure reactor of unknown free volume. This volume was calculated using data on the amount of helium introduced, the pressure, and the temperature prevailing in the reactor. The results of the measurements and calculations are shown in Table 4. The temperature and pressure values recorded in the pressure reactor, immediately after the introduction of helium, were used for the calculations.

**Table 4.** Measurement of the free volume of the reactor based on the amount of helium taken from the gas cylinder and introduced into the reactor without coal.

Test No.	Pressure Inside the Reactor	Average Temperature Inside the Reactor	Amount of Helium Inserted to the Reactor	Free Volume in the Reactor (under Standard Conditions)
	bar	°C	mol	Ndm <sup>3</sup>
1	3.0	11.7	235.23	1857.0
2	5.0	10.5	382.16	1802.6



The difference in the resulting free volume of 54.4 dm<sup>3</sup> is relatively small and is probably due to measurement errors in pressure and temperature. The volume of free space inside the reactor can also be calculated by measuring the amount of helium released from the reactor to the outside. These volumes were measured with a gas meter, and, under the measured conditions, they were 5807 dm<sup>3</sup> and 9342 dm<sup>3</sup> for the No. 1 and No. 2 tests, respectively. The results, converted to standard conditions, are shown in Table 5.

**Table 5.** Measurement of the free volume of the reactor based on the amount of helium released into the atmosphere from the reactor without coal.

Reactor Final Pressure	Atmospheric Pressure	Ambient Temperature	Amount of Helium Released from the Reactor		Free Volume in the Reactor (under Standard Conditions)
			dm <sup>3</sup>	mol	
bar	hPa	°C	dm <sup>3</sup>	mol	Ndm <sup>3</sup>
3.0	965	15.1	5807	230.82	1723.5
5.0	975	1.5	9342	377.27	1690.2

The obtained results of the free volume calculated by this method are 1724 Ndm<sup>3</sup> and 1690 Ndm<sup>3</sup>. They are slightly lower than the values calculated by the previous method. In order to obtain the value closest to the real value, the average value of all these four measurements was calculated. The averaged value of the free volume of the reactor without the inserted coal blocks was 1768.3 Ndm<sup>3</sup>.

To calculate the free volume inside the reactor after inserting the coal blocks, the coal blocks' volume of 458 dm<sup>3</sup> was subtracted from the reactor's total volume of 1768.3 Ndm<sup>3</sup>. The value sought was therefore ~1319 dm<sup>3</sup>.

### 3.1.2. Measurement of the Free Volume of the Reactor with the Coal Block Located Inside

Measurements of the free volume in the reactor were also carried out after the coal blocks were placed in the reactor. The measurement procedure was the same as for the empty reactor. The results of the tests are shown in Tables 6–8.

**Table 6.** Calculation of helium taken from gas cylinders.

Test No.	Initial Pressure in the Helium Cylinder	Final Pressure in the Helium Cylinder	Helium Cylinder Volume	Helium Cylinder Temperature	Amount of Helium Taken from the Gas Cylinder	
	bar	bar			dm <sup>3</sup>	°C
3 (3 bar)	92	11	50	11.3	171.25	3836
4 (6 bar)	187	28	50	11.9	335.44	7514

From the data shown in Table 6, the volume of helium taken for testing at 3 and 6 bar is 3836 Ndm<sup>3</sup> and 7514 Ndm<sup>3</sup>, respectively, under standard conditions. The gas taken from the helium cylinder was introduced into a pressure reactor with coal blocks inserted. The free volume in the pressure reactor was calculated analogously to the measurements without the inserted coal blocks. The results obtained are shown in Table 7.

**Table 7.** Measurement of the free volume of the reactor based on the amount of helium taken from the gas cylinder and introduced into the reactor with coal.

Test No.	Pressure Inside the Reactor	Average Temperature Inside the Reactor	Amount of Helium Inserted to the Reactor	Free Volume in the Reactor (under Standard Conditions)
	bar	°C	mol	Ndm <sup>3</sup>
3	3.0	8.5	171.25	1337
4	6.0	8.7	335.44	1310

**Table 8.** Measurement of the free volume of the reactor based on the amount of helium released into the atmosphere from the reactor with coal.

Reactor Final Pressure	Atmospheric Pressure	Ambient Temperature	Volume of Helium Released from the Reactor		Free Volume in the Reactor (under Standard Conditions)
bar	hPa	°C	dm <sup>3</sup>	Ndm <sup>3</sup>	Ndm <sup>3</sup>
2.95	1018	10.6	3861	3733	1324
5.89	1025	11.2	7442	7235	1285

The difference in the resulting free volumes of 27 dm<sup>3</sup> is relatively low and is probably due to measurement errors in pressure and temperature measurements. As before, the volume of free space inside the reactor was calculated by the second method by measuring the amount of helium released from the reactor. These volumes under measured conditions were 3861 dm<sup>3</sup> and 7442 dm<sup>3</sup> for test No. 3 and test No. 4, respectively. The obtained results converted to standard conditions are shown in Table 8.

The obtained free volume results are close to each other, with a difference of 39 Ndm<sup>3</sup>. It should be noted that there are slight differences between the data shown in Tables 7 and 8. The average value of the free volume of the reactor, calculated based on four measurements with coal blocks located inside, is 1314 dm<sup>3</sup>.

### 3.1.3. Summary of Reactor Free-Volume Measurement Results

Measurements of the reactor's free volume based on tests with an empty reactor and a reactor filled with coal show minimal differences. The empty reactor had a free volume of 1319 dm<sup>3</sup>, while after inserting the coal, the volume decreased to 1314 dm<sup>3</sup>. For further calculations, it was assumed that the free volume of the reactor is 1319 dm<sup>3</sup>.

Based on this, it is possible to estimate the maximum measurement error. From the calculations conducted, the maximum volume deviation is determined is 34 dm<sup>3</sup> (Table 8). In percentage terms, the measurement error is not significant and is approximately 2.6%.

### 3.2. CO<sub>2</sub> Sorption Test

In the next stage of the work, the pressurized reactor with coal placed inside was filled with carbon dioxide to the final pressure of 30.00 bar. The filling time was 7 h and 50 min. Three cylinders of liquid CO<sub>2</sub> were used for this purpose. Two days were allowed to pass while a gradual drop in CO<sub>2</sub> gas pressure was observed, indicating CO<sub>2</sub> sorption by the coal. When the carbon dioxide level decreased, it was recharged to a pressure of 30 bar, and the system was left until the pressure dropped again. The carbon dioxide replenishment process was repeated a total of three times. Then, after a sorption test lasting total 240 h, CO<sub>2</sub> was slowly released from the reactor into the atmosphere. During the test, graphs of pressure and temperature changes were plotted. The relations obtained are presented in Figures 5 and 6.

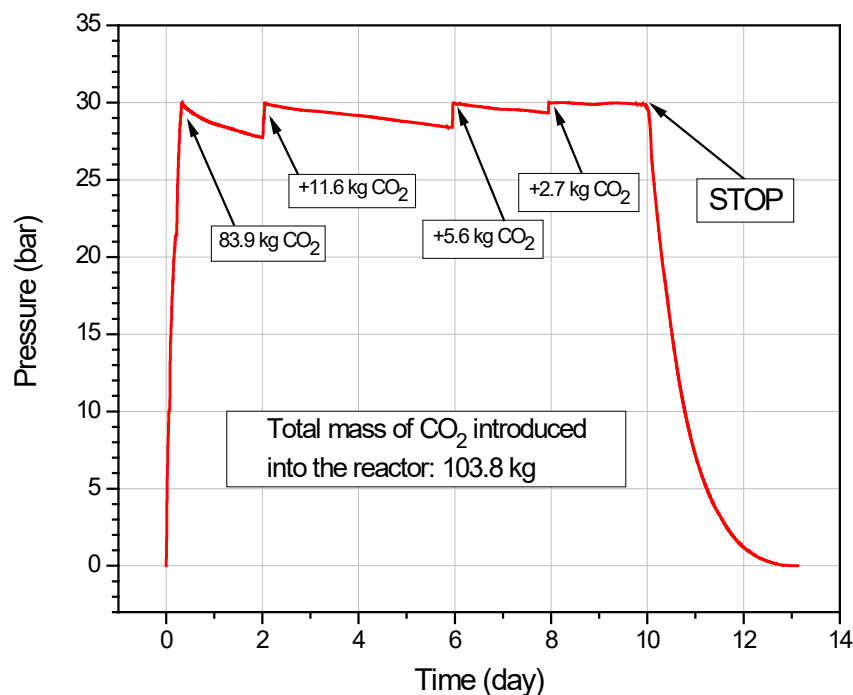


Figure 5. Pressure changes inside the reactor as a function of time during CO<sub>2</sub> sorption and desorption.

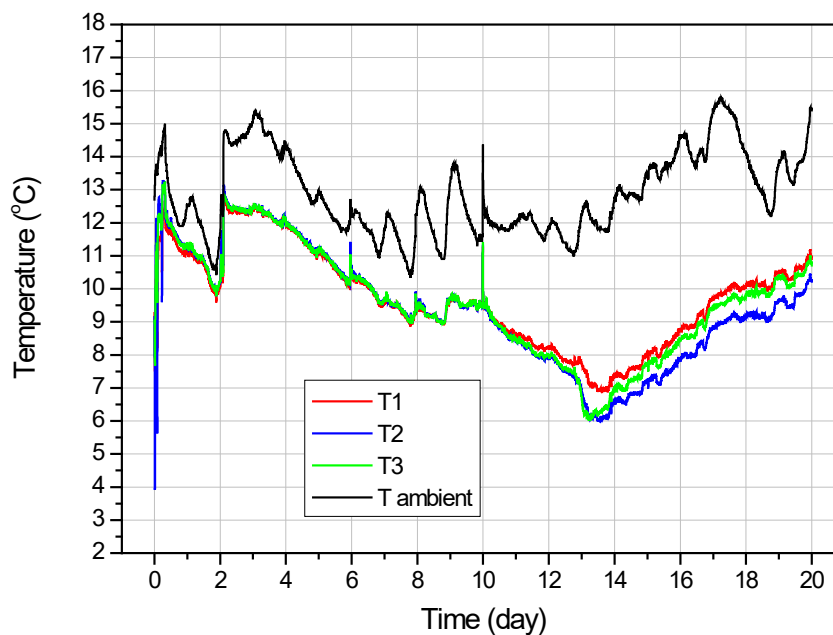


Figure 6. Temperature variation inside the reactor as a function of time.

It can be seen from Figure 5 that, with each successive addition of CO<sub>2</sub> gas, there was a decreasing trend in the reactor's pressure drop, and the volume of CO<sub>2</sub> introduced during each cycle also reduced. The total amount of CO<sub>2</sub> introduced into the reactor was 103.8 kg. The total sorption time was 10 days. After the outlet valves were opened, the time for the pressure in the reactor to drop to atmospheric was 3 days. The amount of CO<sub>2</sub> released from the reactor was measured with a gas meter at all times. Despite the fact that the pressure in the reactor equalized to atmospheric levels, carbon dioxide continued to escape from the reactor, as indicated by increasing values on the gas meter. It was determined that desorption would end if three consecutive CO<sub>2</sub> volume measurements, taken approximately every 8 h at the exit from the reactor, showed no increase in the

amount of desorbed gas. This time was 7 days. The total time for carbon dioxide to be released from the reactor was therefore  $3 + 7 = 10$  days.

From the graph in Figure 6, it can be seen that the temperatures T1–T3 for the duration of the test are very close to each other and their changes are proportional to changes in ambient temperature. It can be seen that the temperature in the reactor increased with each cycle of carbon dioxide introduction. This occurred because the ambient temperature and the temperature of the introduced CO<sub>2</sub> were higher than the temperature inside the pressurized reactor. The apparent fluctuation in ambient temperature is due to the 24-hour cycle of heating up the air during the day and cooling down at night. The average temperature inside the reactor during CO<sub>2</sub> sorption test was 11.6 °C.

#### Calculations of the Sorption Capacity of Coal in Relation to the Sorption of Gaseous CO<sub>2</sub>

Based on knowledge of the total volume of CO<sub>2</sub> introduced into the reactor and the free volume of the reactor (which should be subtracted), the amount of carbon dioxide adsorbed by the coal seam was calculated. The relevant calculations are shown below.

The total amount of carbon dioxide introduced into the reactor, as measured by the mass change of three cylinders of this gas, was 103.8 kg of CO<sub>2</sub>. The amount of carbon dioxide (in moles) that filled the free volume of 1319 Ndm<sup>3</sup> at a pressure of 30.00 bar was calculated from the transformed Equation (2). After substituting the measured data into Equation (2), the amount of CO<sub>2</sub> filling the free volume was 1671.36 moles. Upon converting 1671.36 moles to the mass of CO<sub>2</sub>, a value of 73.54 kg is obtained. By subtracting the 73.54 kg of CO<sub>2</sub> that occupies the free volume in the reactor from the 103.8 kg of CO<sub>2</sub> introduced, the amount of CO<sub>2</sub> adsorbed by the coal can be determined. This value is 30.26 kg. In percentage terms, relative to 631 kg of raw coal, the amount of CO<sub>2</sub> adsorbed was 4.8 wt%.

Figure 7 shows the relationship between time and the volume of pressurized carbon dioxide released from the reactor into the atmosphere. The measurements continued from the opening of the reactor outlet valves until the gas meter stopped indicating a volume increase. This time lasted 10 days.

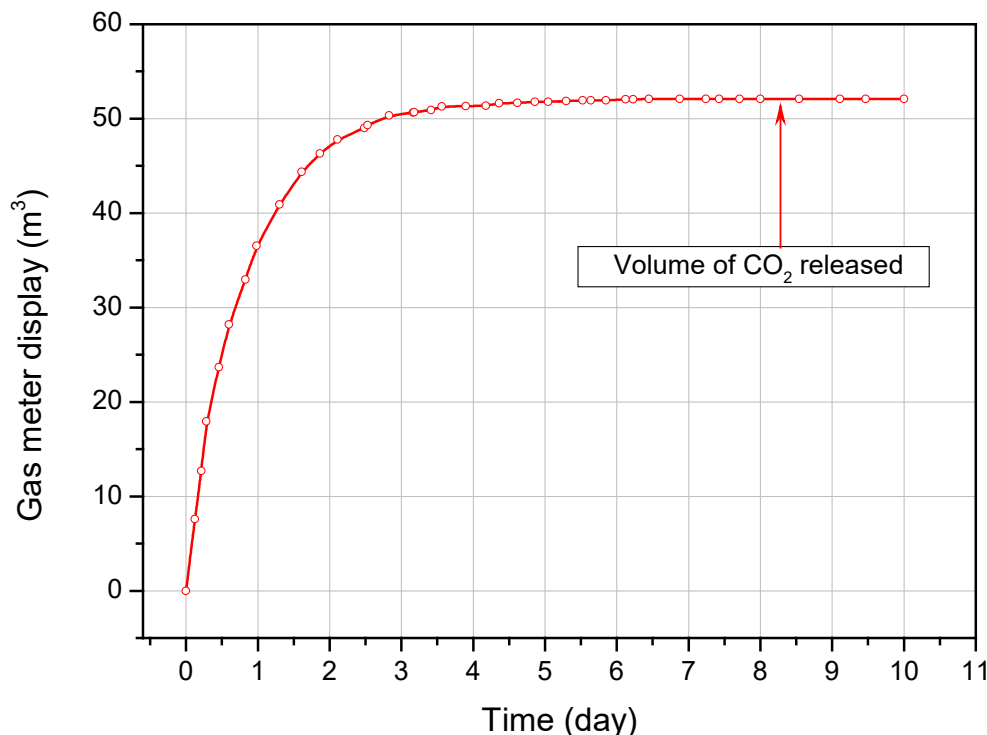


Figure 7. The volume of CO<sub>2</sub> released from the reactor as a function of time.



The average ambient air temperature at the time calculated from the data in Figure 6 was 13.1 °C, while the average ambient pressure was 954 hPa (0.942 bar). The total amount of CO<sub>2</sub> discharged from the reactor under the measured conditions was 52.094 dm<sup>3</sup>, which—converted to standard conditions—gives a value of 46.186 Ndm<sup>3</sup>. Since 52.844 Ndm<sup>3</sup> of CO<sub>2</sub> (103.8 kg) was introduced into the reactor, this means that 6.658 Nm<sup>3</sup> of CO<sub>2</sub> (13.1 kg) remained adsorbed in the coal blocks. In relation to the total amount of adsorbed CO<sub>2</sub> (30.26 kg), this value represents 43.3% of its mass.

#### 4. Discussion and Conclusions

- Ex situ experiments allow us to assess the sorption capacity of coal. Such information allows us to quantify CO<sub>2</sub> storage capacity when designing commercial-scale installations. For coal from the Polish coal mine “Piast-Ziemowit”, the sorption capacity measured at a pressure of 30 bar was 4.8% by weight relative to the raw coal mass.
- In the research study conducted, CO<sub>2</sub> injection followed a multi-stage process where the vessel was repeatedly refilled with gas to achieve a pressure of 30 bar. With each successive CO<sub>2</sub> addition, there was a diminishing trend in the reactor’s pressure drop. Concurrently, the amount of CO<sub>2</sub> introduced during each cycle decreased. This suggests that, as CO<sub>2</sub> migrated and was adsorbed into the coal seam pores, the available sorption capacity progressively diminished.
- Ex situ tests allow simulation of desorption processes following the adsorption phase. Such experiments performed on taken coal samples are essential to determine the total amount of CO<sub>2</sub> that remains adsorbed in the coal even after lowering the pressure. This information is crucial, and the method is recommended for risk management. In the event of an emergency leading to leakage, this allows for an estimation of how much of the injected CO<sub>2</sub> would remain within the coal seam.
- Based on the known geometry of the artificial seam, it is possible to assess the dynamics of CO<sub>2</sub> migration. This knowledge can be transferred to in situ conditions and may prove helpful during an assessment of the feasible injection rates when designing commercial installations.

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