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Crystal habit modification of Cu(II) isonicotinate-*N*-oxide complexes using gel phase crystallisation

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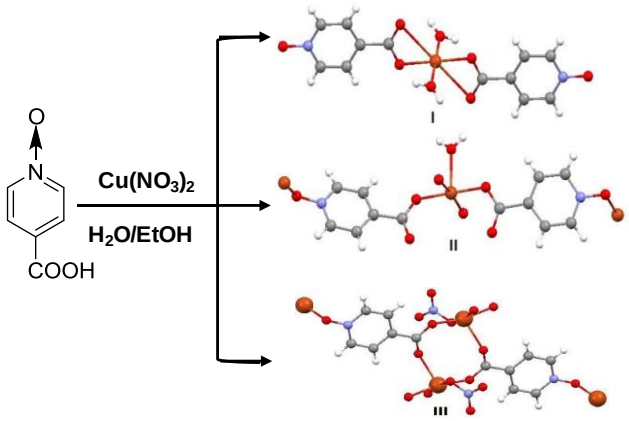
Dipankar Ghosh,[a] Katja Ferfolja,[a] Žygimantas Drabavičius,[a] Jonathan W. Steed [b]* and Krishna K. Damodaran[a]*

We report the crystallisation of three forms of copper(II) isonicotinate-*N*-oxide complex and their phase interconversion via solvent-mediated crystal to crystal transformation. The different forms of copper complex have been isolated and characterised by single crystal X-ray diffraction. Gel phase crystallisation performed in hydrogels, low molecular weight gels and gels of tailored gelator showed crystal habit modification. Crystallization in aqueous ethanol resulted in concomitant formation of blue (form-I) and green (form-II/IV) crystals while use of low molecular weight gels results in selective crystallization of the blue form-I under identical conditions. Comparison of gel phase and the solution state crystallisation in various solvent compositions reveals that the blue form-I is the thermodynamically stable form under ambient conditions.

Introduction

Supramolecular gels based on low molecular weight gelators (LMWGs)[1-9] have witnessed a tremendous growth over the last decade due to their emerging potential applications[5-8] such as dynamic behaviour, use as cell growth media, drug delivery and as media to control crystal growth. Gel phase crystallisation in hydrogels is a classical technique for inorganic compounds and biomolecules such as proteins.[11-13] The gel environment can influence properties[14-20] such as crystal habit, crystal size and polymorphism. Polymorphism depends on a number of factors including nucleation rate, which can give rise to the simultaneous crystallisation of two or more polymorphs with similar nucleation rates, known as concomitant polymorphism.[21] For decades, researchers have been using various techniques such as evaporative, crystallization, solution cooling, melt crystallization and sublimation,[22-23] to search for polymorphic modifications. While these methods are highly effective, they can sometimes fail to efficiently isolate slow-nucleating forms. LMWGs can provide various advantages as crystal growth media because of their versatility, stimuli-responsive properties and often facile synthesis. Gel phase crystallisation results in the shutdown of convection currents leading to diffusion limited growth, and the gel fibres can provide an active surface for heterogeneous nucleation. There have been a few recent reports of crystallisation within gels based on LMWGs[19, 24-27] and small-molecule supramolecular hydrogels have been used to crystallise pharmaceuticals such a modafinil[24] and carbamazepine.[19] An inert gel matrix based on LMWGs (without drug-specific functionality) has been shown to influence pharmaceutical crystallization solid form and habit outcome.[27] Efforts have also been made to develop supramolecular gel phase crystallisation using a gelator that mimics the anticancer drug cisplatin. This resulted in crystal habit modification of cisplatin and the isolation of a novel solvate form.[26] LMWGs that are structurally similar to the crystallisation substrate have been shown to give rise to selective crystallisation of the metastable R polymorph of the highly polymorphic drug precursor ROY.[28] Gel phase crystallisation of isoniazid gives rise to significant differences in crystal habit and crystal size compared to solution control



Scheme 1: Molecular structures of three copper(II) isonicotinate-*N*-oxide complexes deposited in Cambridge Structural Database.[10]

[a] *Department of Chemistry, Science Institute, University of Iceland, Dunhagi 3, 107 Reykjavík, Iceland; Phone: +354 525 4846; Fax: +354 552 8911; E-mail: krishna@hi.is.*
[b] *Department of Chemistry, Durham University, South Road, Durham DH1 3LE, UK; E-mail: jon.steed@durham.ac.uk.*

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experiments.[20] In 2004, Hamilton´s group demonstrated the use of a hydrogel medium to crystallize calcite,[29] while more recently Gunnlaugsson´s group have reported that supramolecular gels can be used to produce single crystal nanowires based on NaCl, KCl, and KI in a gel medium.[30] In the present work, we report the use of LMWGs as crystallisation media for coordination compounds, which display several crystalline forms varying in the copper coordination environment. Specifically, we have selected the complexes of copper(II) with isonicotinic acid-*N*-oxide which exhibits three forms[31-33] deposited within the Cambridge Structural Database[10] and these complexes can be easily isolated (Scheme 1). These three forms of copper(II) isonicotinate-*N*-oxide complex are a discrete square planar diaqua species bound through monodentate carboxylate oxygen atoms $[Cu(C_6H_4NO_3)_2(H_2O)_2]$ (blue, form-I, CSD refcode BUXDED), a square pyramidal aqua complex involving bridging ligands bound by both carboxylate and N-oxide oxygen atoms $[Cu(\mu\text{-}C_6H_4NO_3)_2(H_2O)]_n$ (green, form-II, CSD refcode BEJCID) and a trinuclear hydroxyl-bridged species $[Cu(H_2O)(\mu\text{-}OH)_2\{Cu(C_6H_4NO_3)_2(H_2O)_2\}_2]\cdot 2H_2O$ (green, form-III, CSD refcode BULWIO), formed under basic conditions. A fourth form of formula $[\{Cu(C_6H_4NO_3)_2\}\{Cu(C_6H_4NO_3)NO_3\}_2]_n$ (green, form-IV) is reported herein. The gelators are based on the amide NH···O supramolecular synthon, an important class of stimuli-responsive supramolecular gels[7, 34-50] with tuneable properties. Supramolecular gels based on trimesic amide derivatives[30, 51-53] have been selected as gel media due to their typically very low minimum gel concentration (MGC). Two candidate gels mimic the pyridine N-oxide functional group of the isonicotinic acid-*N*-oxide ligand are reported.

Experimental

Materials and methods

All starting materials were purchased from Sigma Aldrich and were used as supplied. The tris-amide of trimesic acid with L-valine methyl ester (**Val-TMA**)[54] and aminopyridines[51] were synthesized following reported procedures. The tailored N-oxide compound was synthesized by oxidizing the pyridyl group of *tris*-pyridyl trimesic amides.[51] Deionized water was used for all the experiments and absolute ethanol was obtained by distillation over Mg turnings and iodine.[55] ^{1}H-NMR spectra were recorded on a Bruker Advance 400 spectrometer. Single crystal X-ray diffraction (SCXRD) was performed on Bruker D8 venture, powder X-ray (PXRD) was carried out using Bruker D8 Focus instrument. The morphologies of the xerogels were analysed by Scanning Electron Microscopy (SEM) on a Leo Supra 25 Microscope.

Synthesis

*3,3',3''-((benzene-1,3,5-tricarbonyl)tris(azanediyl))tris-(pyridine 1-oxide) (**L-3Nox**)*: To a solution of N^1,N^3,N^5-tri(pyridin-3-yl)benzene-1,3,5-tricarboxamide (0.877 g, 2.00 mmol) in hot methanol (50 mL), *m*-chloroperoxybenzoic acid (1.86 g, 10.8 mmol) was added in portions over a period of 15 minutes. The reaction mixture was refluxed at 70 °C overnight. The mixture was cooled to room temperature. The solid obtained was filtered, washed with water followed by methanol, dried and the product was obtained as white powder (0.76 g, 1.56 mmol). Yield: 78%. ^{1}H NMR (400 MHz, DMSO-d_6) δ: 10.95 (3H, s), 8.86 (3H, s), 8.74 (3H, s), 8.05 (3H, d, J = 6.4), 7.71 (3H, d, J=8.4), 7.46 (3H, dd, J= 8.6, 6.2). HRMS (m/z) Calcd. for $C_{24}H_{18}N_6O_6Na$: 509.118; found: 509.118 [M-Na^+]. Anal. data for $C_{24}H_{18}N_6O_6$: Calc. C, 59.26; H, 3.73; N, 17.28. Found: C, 59.10; H, 3.82; N, 17.00.

$[Cu(C_6H_4NO_3)_2(H_2O)_2]$ (form-I): 27.8 mg (0.2 mmol) of isonicotinic acid-*N*-oxide was dissolved in water (8.5 mL)/ ethanol (1.5 mL) mixture and was layered over 5 mL aqueous solution of 24.1 mg $Cu(NO_3)_2\cdot 3H_2O$ (0.1 mmol). Slow evaporation of the mixture resulted blue crystals of form-I in 3-4 days.[31]

$[Cu(\mu\text{-}C_6H_4NO_3)_2(H_2O)]_n$ (form-II): 27.8 mg (0.2 mmol) of isonicotinic acid-*N*-oxide was dissolved in ethanol (10 mL) and layered over ethanolic solution (5 mL) of 24.1 mg $Cu(NO_3)_2\cdot 3H_2O$ (0.1 mmol) and the vial was sealed. Plate shaped green crystals of form-II were obtained in 3-4 days.[33]

$[\{Cu(C_6H_4NO_3)_2\}\{Cu(C_6H_4NO_3)NO_3\}_2]_n$ (form-IV): 27.8 mg (0.2 mmol) of isonicotinic acid-*N*-oxide and 24.1 mg (0.1 mmol) of $Cu(NO_3)_2\cdot 3H_2O$ was added to 2 mL of ethanol and the mixture was heated at 85 °C in a sealed vial. Block shaped green crystals of form-IV were obtained overnight at 85 °C. Anal. data for $C_{24}H_{16}Cu_3N_6O_{18}$: Calc. C, 33.25; H, 1.86; N, 9.69. Found: C, 33.16; H, 1.94; N, 9.70.

Single Crystal X-ray Diffraction

X-ray quality single crystals of form-IV isolated from mother liquor, immediately immersed in cryogenic oil and then mounted. The diffractions were collected using MoKα radiation (λ =0.71073Å) on a Bruker D8 Venture (Photon100 CMOS detector) diffractometer equipped with a Cryostream (Oxford Cryosystems) open-flow nitrogen cryostats at the temperature 150.0(2) K. The unit cell determination, data collection, data reduction, structure solution/refinement and empirical absorption correction (SADABS) were carried out using Apex-III (Bruker AXS: Madison, WI, 2015). The structure was solved by direct method and refined by full-matrix least squares on F^2 for all data using SHELXTL[56] and Olex2[57] software. All non-disordered non-hydrogen atoms were refined anisotropically except for the disordered oxygen atom of the nitrate group, where the free variables were refined by FVAR instruction. All the hydrogen atoms were placed in the calculated positions and refined using a riding model.

Gelation property of tailored N-oxide compound

The gelation ability of **L-3Nox** was screened in various solvent systems (see ESI). The gelator was soluble only in high polar solvents such as DMF, DMA and DMSO and gels were formed only DMSO/water mixtures. In a typical experiment, the

compound was dissolved in required amount of DMSO by heating and sonicating followed by the addition of water. The mixture was then sonicated to form a suspension, left undisturbed to form the gel. Gel formation was confirmed by the inversion test.

Minimum gelation concentration: Various amounts (1.0 to 5.0 mg) of **L-3Nox** gelator were taken in a standard 7 mL vials and 0.5 mL of DMSO was added. The solution was heated and sonicated to dissolve the compound followed by the addition of 0.5 mL water and the resulting mixture was left undisturbed to form the gel. After 24 hours, gel formation was checked by inversion test. The lowest concentration at which gel was formed was recorded as minimum gel concentration (MGC).

Gel-sol transition temperature: The required amount of **L-3Nox** gelator was taken in a standard 7 mL vial and the gel was prepared in DMSO/water (1:1 v/v) as per the above procedure. After 24 hours, a small spherical glass ball weighing 92 mg was placed on the gel surface, and the gel was heated gradually in an oil bath. The temperature at which the ball touched the bottom of the vial was recorded as the gel-sol transition temperature (T_{gel}).

Crystallisation experiments

Crystal to crystal transformation: Transformation of blue complex (form-I) to green polymer (form-II): Form-I (3mg) was added to absolute ethanol (10 mL) in a standard vial. The vial was sealed to prevent the evaporation of ethanol and left without disturbance. Green crystals of form-II were obtained over the course of one week and were characterized by SCXRD.

Transformation from green polymer (form-II) to blue complex (form-I): Water (1 mL) was added to form-II (3 mg). Excess water was avoided since form-II is sparingly soluble in water. The transparent green crystals almost instantly turned into opaque blue material and the mixture was left undisturbed. X-ray quality blue crystals of form-I were obtained in 2-3 days, and were characterized by SCXRD.
Solid state transformation from green polymer (form-IV) to blue complex (form-I): Water (1 mL) was added to form-IV (3 mg) in a standard vial. X-ray quality blue crystals were obtained in 2-3 days and characterized by SCXRD.

Scanning Electron Microscopy: The gelator (**L-3Nox**) was dissolved in 0.5 mL of DMSO, the mixture was heated and sonicated, 0.5 mL of water was added to form the gel. The gel was filtered through a filter paper after 24 hours and the residue was air dried in fume hood. The xerogels were gold coated and SEM was performed on a Leo Supra 25 Microscope.

Gel Phase Crystallisation

Crystallisation of $Cu(NO_3)_2.3H_2O$ and isonicotinic acid-*N*-oxide was performed in presence of hydrogelators (agarose and gelatin) and low molecular weight gelators (LMWGs). In a typical experiment, isonicotinic acid-*N*-oxide (2 equivalents) and the gelator were dissolved together in water (for agarose and gelatin) and then $Cu(NO_3)_2.3H_2O$ (1 equivalent) was added. The solution was sonicated and left undisturbed to form gel and crystallise. For LMWGs, the isonicotinic acid-*N*-oxide and the gelator were dissolved in a polar organic solvent (DMF or DMSO) and $Cu(NO_3)_2.3H_2O$ and water were added to this mixture in the quantities given below.

Crystallisation in agarose: Isonicotinic acid-*N*-oxide (13.9 mg, 0.1 mmol) and agarose (6 mg) were dissolved in water (1 mL) by heating, and $Cu(NO_3)_2.3H_2O$ (12.1 mg, 0.05 mmol) was added to the resulting solution. X-ray quality crystals of form-I were obtained in 2 days.

Gel phase crystallisation in Val-TMA: isonicotinic acid-*N*-oxide (27.8 mg, 0.2 mmol) and **Val-TMA** (40 mg, 4 wt%) were dissolved in DMF (0.5 mL) by heating and sonicating. Water (0.5 mL) and $Cu(NO_3)_2{\cdot}3H_2O$ (24.1 mg, 0.1 mmol) was added to this solution. The mixture was left undisturbed to form a blue gel. X-ray quality crystals of form-I were obtained in 2-3 days.

Gel phase crystallisation in L-3Nox: isonicotinic acid-*N*-oxide (27.8 mg, 0.2 mmol) and **L-3Nox** (8 mg) were dissolved in DMSO (0.5 mL) by heating and sonicating. Water (0.5 mL) and $Cu(NO_3)_2{\cdot}3H_2O$ (24.1 mg, 0.1 mmol) were added to this solution and leaving the mixture undisturbed yielded blueish green gel. X-ray quality single crystals of form-I were obtained in 2 days.

Results and discussion

Solution phase crystallization

Layering an ethanolic solution of $Cu(NO_3)_2{\cdot}3H_2O$ over an aqueous ethanolic solution (5% of water, v/v), of isonicotinic acid-*N*-oxide at a 1:2 metal-ligand ratio results in the formation of a mixture of blue and green crystals over a period of three days (Figure S1). These crystals were isolated and their structures determined by single crystal X-ray diffraction. This technique shows the samples to match the reported structures, namely form-I (blue crystals) and form-II (green crystals).[31, 33] Formation of the hydroxyl-bridged species form-III in not expected under these conditions due to the absence of base.[32] Basic conditions were not investigated because of the effect of changing pH on the gelation process in LMWGs.[58]

We have optimized the crystallisation conditions for these two forms and confirmed that crystallisation depends on the ethanol/water ratio (ESI). Form-I can be obtained in aqueous ethanol containing >10% of water (v/v), whereas the green form-II is formed at lower water content (<3.5 % of water, v/v). Both forms are obtained simultaneously from aqueous ethanol containing 7% water (v/v). The monoaqua form-II transforms into the diaqua form-I over three days, presumably due to absorption of moisture from the mother liquor. The effect of temperature was studied by comparing the room temperature and low temperature (-15 °C) crystallisation for both forms. The yield of the single crystals in both cases were low

compared to room temperature crystallisation. We have also performed low temperature crystallisation for form-I & II and interestingly, form-II does not disappear from the mixture even after three weeks. The conversion of green to blue crystals prompted us to explore solid state transformation of form-I & II. Crystals of form-I were isolated, immersed in

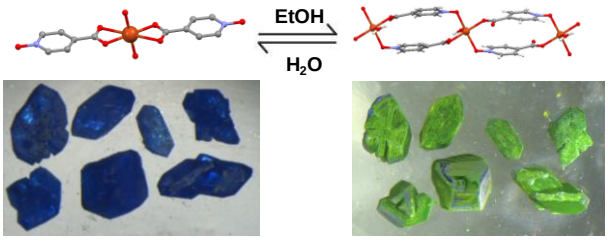


Figure 1: Solvent mediated crystal to crystal transformation of form I & II.

absolute ethanol and over five days the crystals underwent conversion to green form-II (Figure 1). Similarly, treating form-II with water resulted in form-I overnight. The transformation of form-I to form-II was found to be reversible, which was confirmed by single crystal X-ray diffraction.

Heating copper(II) nitrate and isonicotinic acid-*N*-oxide at 85 °C in ethanol in a sealed vial resulted in block shaped green crystals, a morphology that contrasts to the plate shaped form-II. Single crystal X-ray analysis of the block shaped green crystals revealed a new coordination polymer (form-IV) of formula $[\{Cu(C_6H_4NO_3)_2\}\{Cu(C_6H_4NO_3)NO_3\}_2]_n$, which displays two different Cu(II) centres with distorted octahedral and square pyramidal geometries (Figure 2a). The oxygen atoms of the carboxylate moiety of the isonicotinate-*N*-oxide ligands are coordinated to the two Cu(II) centres in a bidentate fashion

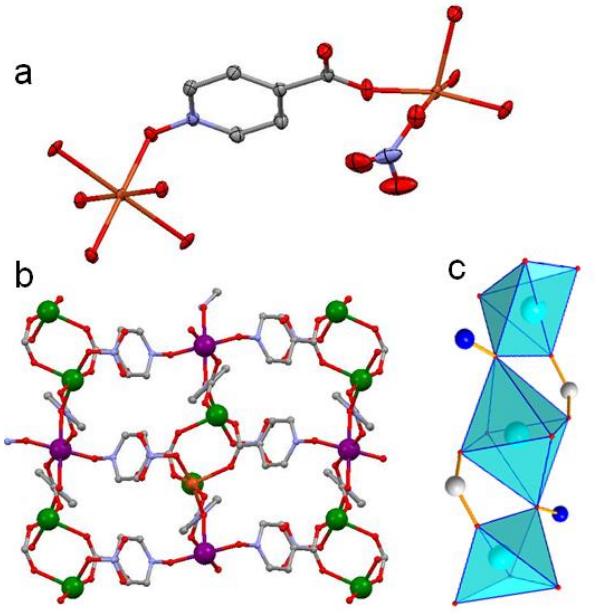
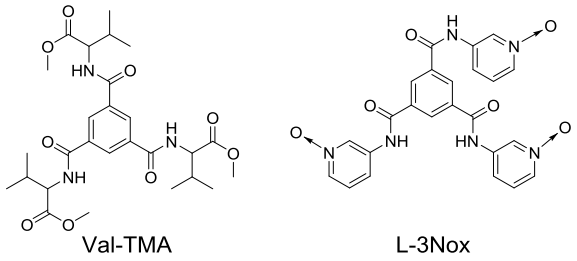


Figure 2: a) Molecular structure of form-IV $[\{Cu(C_6H_4NO_3)_2\}\{Cu(C_6H_4NO_3)NO_3\}_2]_n$ b) representation of crystal structure showing octahedral (purple) and square pyramidal (green) metal centres and c) interconnected octahedral and square pyramidal geometry.

forming an eight membered metallo-macrocycle. The Cu(II) centres in the metallo-macrocycle display square pyramidal geometry with oxygen atoms of N-oxide moiety, nitrate anion and two carboxylate moieties in the equatorial site and the axial position is coordinated to the oxygen atom of the carboxylate moiety.

The oxygen atoms of the N-oxide moieties of isonicotinic acid-*N*-oxide in the metallo-macrocycle are coordinated to the Jahn-Teller distorted octahedral Cu(II) centre. In the octahedral Cu(II) centre, four isonicotinate-*N*-oxide ligands are coordinated to the equatorial position of the metal centre (two oxygen atoms of the N-oxide moiety and two carboxylate oxygen atoms) and the axial positions are coordinated to the oxygen atoms of the N-oxide moiety. The oxygen atom of the N-oxide moiety displays a bridging coordination mode and binds to the equatorial position of the distorted octahedral and square pyramidal copper(II) centres resulting in a complex 3-D network (Figure 2b). Form-IV was found to convert to



Scheme 2: Chemical structure of non-tailored gelator (**Val-TMA**) and tailored N-oxide gelator (**L-3Nox**).

form-I when immersed in water (confirmed by single crystal X-ray diffraction).

Gel phase crystallization

The existence of at least four distinct copper(II) isonicotinate-*N*-oxide complexes prompted us to explore the selective crystallisation of this system in gel media. The copper(II) isonicotinate-*N*-oxide mixtures were initially crystallized in gels of commercially available hydrogelators, namely agarose and gelatin. A blue gel was obtained by dissolving $Cu(NO_3)_2{\cdot}3H_2O$, isonicotinic acid-*N*-oxide and agarose in hot water, which subsequently produced X-ray quality crystals of blue form-I on cooling. Crystallisation experiments performed at various concentrations of agarose, $Cu(NO_3)_2{\cdot}3H_2O$ and isonicotinic acid-*N*-oxide gave similar results, however, experiments with low concentration of metal salt and ligand did not yield any solid product. Crystals of isonicotinic acid-*N*-oxide (CSD refcode XUCPAO) were obtained at higher concentration of metal salt and ligand. Experiments performed with 1:1 ethanol/water (v/v) gave similar results. Use of gelatin hydrogels as the crystalizing medium resulted in blue gels but crystals were not formed even after several weeks. Neither gelatin nor agarose form gels in absolute ethanol. We then turned our attention to LMWGs based on **Val-TMA**[54] (Scheme 2 and ESI).

Gel phase crystallisation was performed by mixing copper(II) nitrate trihydrate and isonicotinic acid-*N*-oxide at 1:2 metal-ligand ratio with the gelator (4.0 wt%) in EtOH/water and DMF/water (1:1, v/v respectively). This resulted in blue

crystals of form-I after two days (confirmed by single crystal X-ray diffraction). The formation of form-I in this 'generic' supramolecular gel medium, prompted us to investigate the crystallisation of copper(II) isonicotinate complexes in tailored gel of LMWGs that are structurally similar to the crystallisation substrate. Recently we have shown that tailored LMWGs gel can enable selective crystallisation[28] of particular polymorphs of the olanzapine precursor, ROY.[59] We have also reported the

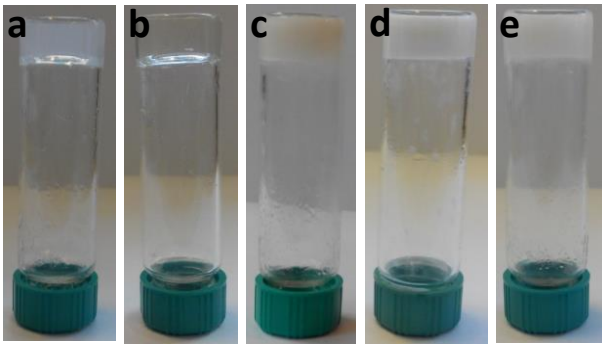


Figure 3: Gels obtained from (a) agarose in water, (b) gelatin in water, (c) Val-TMA in DMF/water (1:1 v/v), (d) Val-TMA in EtOH and (e) **L-3NOx** in DMSO/water (1:1 v/v).

formation of crystalline material in a supramolecular gel matrix of metallogels with pyridyl amides.[60]

The gelator design involves trimesic pyridyl amide based compounds,[7, 34] which show two types of hydrogen bonding motif, one through a N–H···O synthon involving the amide moiety and the other through a N–H···N synthon involving the amide and pyridyl ring nitrogen atom. Thus, tailored gel was designed for the copper(II) isonicotinate-*N*-oxide system by modifying the pyridyl groups of trimesic amides to give *N*-oxides. In this context, we prepared tailored tris-*N*-oxide compound (**L-3Nox**, Scheme 2) by oxidising *N*,*N*′,*N*′′-tris(3-pyridyl)-trimesic amide[51] with *m*-chloroperbenzoic acid (ESI). The gelation properties of **L-3Nox** was tested in aqueous solutions of highly polar DMSO, due to their poor solubility in other solvents. These compounds formed gels in various mixtures of DMSO and water (ESI) and the gel formation was confirmed by inversion test (Figure 3). Although, the optimised conditions for **L-3Nox** was found to be a DMSO/water (1:1, v/v) mixture, we selected a 7:3 (DMSO/water, v/v) mixture for **L-3Nox** since the gel was transparent at this composition. The thermal stability of tailored gel was evaluated by analysing the temperature at which the gel was converted into a liquid phase (T_{gel}). The T_{gel} for **L-3Nox** was found to be 107 °C at 0.5 wt% and minimum gel concentration (MGC) was 0.15 wt% in DMSO/water (7:3, v/v) mixture. The T_{gel} in DMSO/water (1:1 v/v) for **L-3Nox** was 105 °C at 0.5 wt% and MGC was 0.3 wt% in DMSO/water (1:1 v/v) mixture. Moreover, replacing the pyridyl group with pyridyl N-oxide has a significant effect on the MGC of **L-3Nox** (0.15 wt%) compared *to N,N′,N′′*-tris(3-pyridyl)-trimesic amide (0.03 wt%) in DMSO/water (7:3, v/v) mixture.[51]

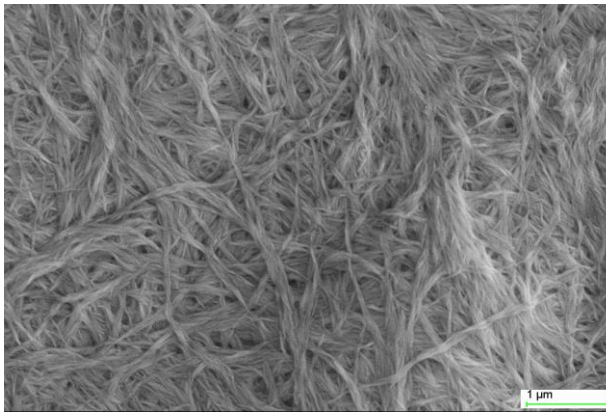


Figure 4: SEM images of xerogels of **L-3Nox** at 0.5 wt% obtained from DMSO/water (7:3, v/v).

Scanning electron microscopy (SEM) was performed on the xerogel of **L-3Nox** to elucidate the morphologies of the xerogels (Figure 4), which clearly indicate the fibrous nature of the gel network and twisted fibres were observed in **L-3Nox**. The thickness of the individual fibers of **L-3Nox** were found to be 20-40 nm and these fibres combines to form bundles with thickness ranging from 100 to 150 nm.

Gel phase crystallisation was performed by mixing copper(II) nitrate trihydrate and isonicotinic acid-*N*-oxide at 1:2 metal-ligand ratio with the **L-3Nox** gelator in DMSO/water. The mixture was heated to give a clear green solution and left without disturbing for 2 hours to form the gel (ESI). The crystallisation experiments were performed at higher concentration of **L-3Nox** gelators (greater than the MGC) to ensure robust gel formation. The crystallisation experiments were performed at varying amounts of copper(II) nitrate trihydrate and isonicotinic acid N-oxide. A greenish-blue gel was formed in an hour in all cases for **L-3Nox** gelator at 0.8 wt% in 1:1 (v/v) DMSO/water. The optimized concentration for crystallisation was found to be 0.058 mmol of copper(II) nitrate trihydrate and 0.116 mmol isonicotinic acid-*N*-oxide. Crystals were not formed for experiments with lower metal-ligand concentration, whereas a concentration of 0.058 mmol or more copper(II) nitrate trihydrate produced blue crystals (form-I) in a week (Figure 5c).

We have compared the morphologies of the form-I crystals obtained from the gel phase crystallisation in different

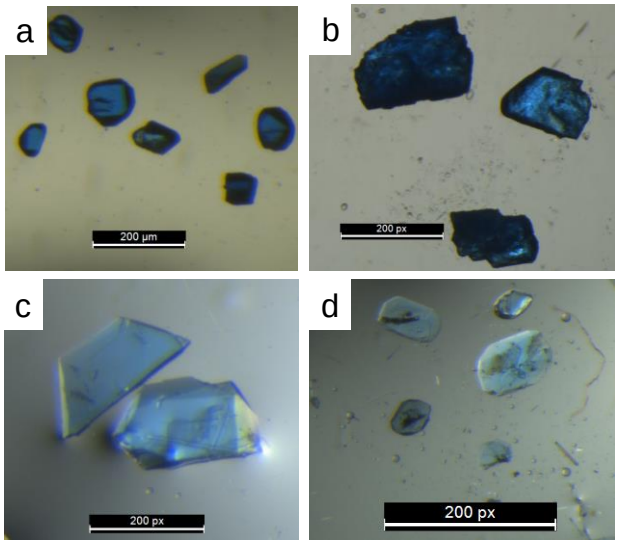


Figure 6: Images of isolated crystals from (a) solution phase, (b) **agarose gel,** (c) **Val-TMA** gel and (d) **L-3Nox** gel.

gelators. The solution phase crystallisation in water/ethanol mixture without gelator resulted in block shaped crystals. Similar crystals of form-I were obtained in water with very low yield due to the crystallisation of the ligand isonicotinic acid-*N*-oxide. The gel phase crystallisation of the complex in agarose gel yielded block shaped crystals (Figure 6). However, plate shaped crystals were obtained from **Val-TMA** and **L-3NOX** tailored gel indicating that the presence of a similar functional groups in the gelator plays a role in crystal morphology.

Crystallisation of copper(II) isonicotinate-*N*-oxide complexes was also performed in ethanolic supramolecular gel media. Since the tailored gelator and agarose do not form gel in absolute ethanol or a mixture of ethanol and other solvents, gel phase crystallisation in ethanol was performed only with **Val-TMA**. Adding copper(II) nitrate to a hot solution of isonicotinic acid N-oxide and **Val-TMA** in absolute ethanol led to a green solution, which subsequently formed a green gel.

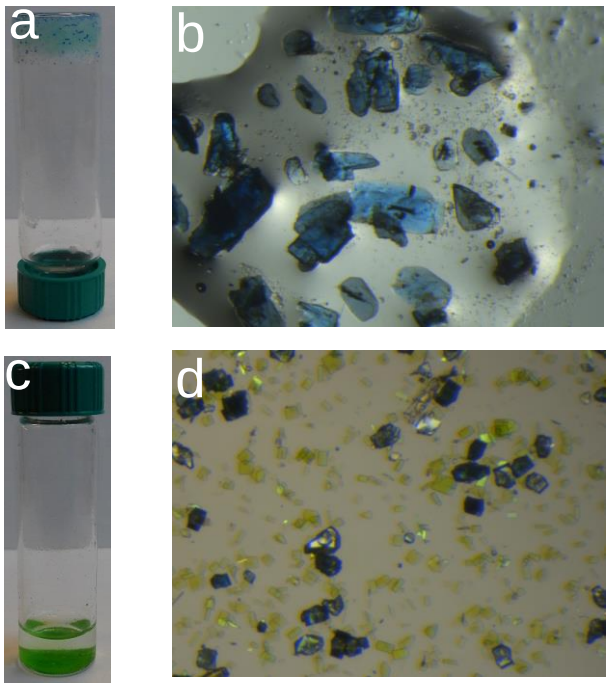


Figure 7: Crystallisation experiments of copper(II) isonicotinate-*N*-oxide complex in water-ethanol (5% water, v/v) (a) with **Val-TMA** gelator (4 wt%) and (b) the isolated crystals. Solution phase crsytalision (c) without gelator and (d) top view of crsytals.

However, crystals of Cu(II)- isonicotinate-*N*-oxide complex were not formed and isonicotinic acid-*N*-oxide crystallised due

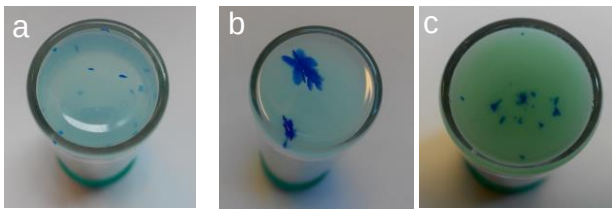


Figure 5: Gel phase crystallisation of Cu(II) complexes from (a) agarose gel in water (1.5 wt%), (b) **Val-TMA** gel at 4.0 wt% in DMF/water (1:1, v/v) and (c) **L-3Nox** gel at 1.0 wt% in DMSO/water (1:1, v/v).

to its poor solubility in ethanol at ambient temperature. The solubility of isonicotinic acid-*N*-oxide was increased by adding trace amount of water. Thus, a mixture of isonicotinic acid N-oxide (0.1 mmol) and **Val-TMA** (4.0 wt%) was heated in water-ethanol (5.0% water, v/v) mixture and copper(II) nitrate (0.1 mmol) was added to yield a green solution and the vial was sealed. The solution turned into a greenish blue gel in 15 minutes, the blue colour intensified overnight resulting in blue crystals of form-I in the gel medium (Figure 7a, b). Blue crystals of form-I were formed in almost every case, with one out of 25 trails formed green crystals of form-II, which might be due to accidental heteroseeding.

Experiments were performed under identical conditions (same solvent composition and same concentration of copper(II) nitrate and isonicotinic acid N-oxide) without **Val-TMA** by adding copper(II) nitrate to a hot solution of isonicotinic acid N-oxide in water-ethanol (5.0 % water, v/v), which resulted in a green solution. The vial was sealed and a mixture of blue and green crystals were formed overnight (Figure 7c, d). X-ray single crystal diffraction of these crystals revealed that blue crystal belongs to form-I and the green crystals were either form-II or form-IV. Thus, concomitant crystallisation of different forms was observed from solution. Most of the green crystals eventually turn blue over a week.

Conclusions

In summary, we report the crystallisation of three forms of copper(II) isonicotinate-*N*-oxide (form-I, II and IV) and their solvent-mediated interconversion. The transformation of a concomitant mixture of blue and green crystals were also studied as a function of solvent concentration. We have designed gelator that is structurally similar to the crystallisation substrate. Gel phase crystallisation of the complex were performed in hydrogels, gels of low molecular weight gelators and tailored LMWG. The morphologies of crystals obtained from solution and from agarose gel proved to be similar to one another while gel phase crystallisation performed in LMWGs and tailored gelator resulted in plate shaped crystals indicating an influence of the gelator on crystallization process. Crystallisation of copper(II) isonicotinate-N-oxide complexes in aqueous ethanol (5.0 % water, v/v) resulted in a mixture of blue and green crystals

whereas gel phase crystallisation in **Val-TMA** gel under identical conditions resulted in only blue crystals indicating the influence of LMWGs in selective crystallisation of the thermodynamically stable form-I.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

‡ Crystal data for Cu(II)-isonicotinic acid N-oxide complex (form-IV): $C_{24}H_{16}Cu_3N_6O_{18}$, FW = 867.05, Monoclinic, $P2_1/n$, a = 8.892(2), b =15.096(4), c = 11.178(3) Å, β = 103.646(9)°, V = 3443(3) Å³, Z = 4, D_c = 1.975 cm³, *F(000)* = 866, T = 150 K. Final residuals (for 242 parameters) were *R1* = 0.0218 for 3288 reflections with $I > 2\sigma(I)$, and R_1 = 0.0218, wR_2 = 0.0581, GOF = 0.989 for all 3052 reflections. CCDC 1846767 contains the supplementary crystallographic data for this paper. These data can be obtained free of charge via www.ccdc.cam.ac.uk/data_request/cif, or by emailing data_request@ccdc.cam.ac.uk, or by contacting The Cambridge Crystallographic Data Centre, 12 Union Road, Cambridge CB21EZ, UK; fax: +44 1223 336033. See http://dx.doi.org/10.1039/xxx for other electronic format.

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