

## Article

# Tuning Photophysical Properties by *p*-Functional Groups in Zn(II) and Cd(II) Complexes with Piperonylic Acid

Francisco Sánchez-Férez <sup>1</sup>, Joaquim M<sup>a</sup> Rius-Bartra <sup>1</sup>, José A. Ayllón <sup>1</sup>, Teresa Calvet <sup>2</sup>, Mercè Font-Bardia <sup>3</sup> and Josefina Pons <sup>1,\*</sup>

<sup>1</sup> Departament de Química, Universitat Autònoma de Barcelona, 08193 Barcelona, Spain; francisco.sanchez.ferez@uab.cat (F.S.-F.); joaquim.rius.bartra@gmail.com (J.M.R.-B.); joseantonio.ayllon@uab.cat (J.A.A.)

<sup>2</sup> Departament de Mineralogia, Petrologia i Geologia Aplicada, Universitat de Barcelona, Martí i Franquès s/n, 08028 Barcelona, Spain; mtcalvet@ub.edu

<sup>3</sup> Unitat de Difracció de Raig-X, Centres Científics i Tecnològics de la Universitat de Barcelona (CCiTUB), Universitat de Barcelona, Solé i Sabarís, 1-3, 08028 Barcelona, Spain; mercef@ccit.ub.edu

\* Correspondence: josefina.pons@uab.cat; Tel.: +34-935812895

**Abstract:** Aggregation between discrete molecules is an essential factor to prevent aggregation-caused quenching (ACQ). Indeed, functional groups capable of generating strong hydrogen bonds are likely to assemble and cause ACQ and photoinduced electron transfer processes. Thus, it is possible to compare absorption and emission properties by incorporating two ligands with a different *bias* toward intra- and intermolecular interactions that can induce a specific structural arrangement. In parallel, the  $\pi$  electron-donor or electron-withdrawing character of the functional groups could modify the Highest Occupied Molecular Orbital (HOMO)–Lowest Unoccupied Molecular Orbital (LUMO) energy gap. Reactions of  $M(\text{OAc})_2 \cdot 2\text{H}_2\text{O}$  ( $M = \text{Zn}(\text{II})$  and  $\text{Cd}(\text{II})$ ;  $\text{OAc} = \text{acetate}$ ) with 1,3-benzodioxole-5-carboxylic acid (Piperonylic acid, HPip) and 4-acetylpyridine (4-AcPy) or isonicotinamide (Isn) resulted in the formation of four complexes. The elucidation of their crystal structure showed the formation of one paddle-wheel  $[\text{Zn}(\mu\text{-Pip})_2(4\text{-AcPy})]_2$  (**1**); a mixture of one dimer and two monomers  $[\text{Zn}(\mu\text{-Pip})(\text{Pip})(\text{Isn})_2]_2 \cdot 2[\text{Zn}(\text{Pip})_2(\text{HPip})(\text{Isn})] \cdot 2\text{MeOH}$  (**2**); and two dimers  $[\text{Cd}(\mu\text{-Pip})(\text{Pip})(4\text{-AcPy})]_2$  (**3**) and  $[\text{Cd}(\mu\text{-Pip})(\text{Pip})(\text{Isn})_2]_2 \cdot \text{MeOH}$  (**4**). They exhibit bridged (**1**,  $\mu_2\text{-}\eta^1\text{:}\eta^1$ ), bridged, chelated and monodentated (**2**,  $\mu_2\text{-}\eta^1\text{:}\eta^1$ ,  $\mu_1\text{-}\eta^1\text{:}\eta^1$  and  $\mu_1\text{-}\eta^1$ ), or simultaneously bridged and chelated (**3** and **4**,  $\mu_2\text{-}\eta^2\text{:}\eta^1$ ) coordination modes. Zn(II) centers accommodate coordination numbers 5 and 6, whereas Cd(II) presents coordination number 7. We have related their photophysical properties and fluorescence quantum yields with their geometric variations and interactions supported by TD-DFT calculations.

**Keywords:** Zn(II) and Cd(II); 4-acetylpyridine; isonicotinamide; piperonylic acid; photoluminescence



**Citation:** Sánchez-Férez, F.; Rius-Bartra, J.M.; Ayllón, J.A.; Calvet, T.; Font-Bardia, M.; Pons, J. Tuning Photophysical Properties by *p*-Functional Groups in Zn(II) and Cd(II) Complexes with Piperonylic Acid. *Molecules* **2022**, *27*, 1365. <https://doi.org/10.3390/molecules27041365>

Academic Editors: Mário J. F. Calvete and Mariette M. Pereira

Received: 19 January 2022

Accepted: 15 February 2022

Published: 17 February 2022

**Publisher's Note:** MDPI stays neutral with regard to jurisdictional claims in published maps and institutional affiliations.



**Copyright:** © 2022 by the authors. Licensee MDPI, Basel, Switzerland. This article is an open access article distributed under the terms and conditions of the Creative Commons Attribution (CC BY) license (<https://creativecommons.org/licenses/by/4.0/>).

## 1. Introduction

The correlation between composition, structure and properties has always been the foundation of materials design. Understanding this interplay allows to comprehend the structure–properties relationship and make significant progress in fields *inter alia* luminescence [1], sensing [2] or photoelectrical conductivity [3,4], allowing to design prominent materials. In the field of optical materials, the conception of fluorescent light emitting materials based on discrete molecular complexes provides a crucial advantage compared to their polymeric analogues as they have better solvent processability [5].

Pyridine based fluorophores have been developed as a fitting family of ligands with which to design fluorescent complexes. They excel at being sensitive to electronic perturbation and are capable of coordinating both soft and hard metal ions [6]. However, they commonly present fluorescence quenching associated with intramolecular charge transfer (ICT) and intra-ligand charge transfer (ILCT) transitions, photoinduced electron

transfer (PET) processes, or aggregation-caused quenching (ACQ) by exciplex formation that promote non-radiative decays. In this scenario, a well-established strategy is to incorporate transition metal ions that can stabilize the excited state as well as minimize PET processes by fixing non-bonding electrons on the heteroatoms of the ligand and lowering the energy of the lone pair orbitals. Diamagnetic metal ions as Zn(II) and Cd(II) are able to coordinate with the fluorophore and present outstanding performance by minimizing relaxation via non-radiative decays [7,8]. Their fully populated d orbitals do not partake in electronic transitions and, therefore, ligand centered  $\pi \rightarrow \pi^*$  transitions are favored, which are desirable in the seeking of high emissive complexes [9]. In particular, being coordinated with *N*- and *O*-donor ligands, they tend to construct highly blue emissive materials [10].

In terms of their metal preferences, Zn(II) presents coordination numbers from 3 to 6, whereas Cd(II) can accommodate up to 8. Their geometries range from tetrahedral to octahedral or, in the case of Cd(II), the less common square-antiprism [11]. In complexes containing carboxylic acids and pyridine derivative ligands, coordination numbers 4–6 are preferred by Zn(II), while Cd(II) is likely to present 6 and 7. The Zn(II) center tends to form dimeric motifs as  $[\text{Zn}_2(\text{CO}_2)_2]$  with two carboxylates coordinated in a bridging mode or square pyramidal  $[\text{Zn}_2(\text{CO}_2)_4]$  paddle-wheels with two apical *N*-donors. Conversely, Cd(II) is likely to construct double bridged  $[\text{Cd}_2(\text{CO}_2)_2]$  cores with four *N*-donor ligands, being the heptacoordinated environment is completed by chelate carboxylates and by designing a  $[\text{Cd}_2(\text{CO}_2)_4]-2+2$  motif. This core is commonly hindered for Zn(II) complexes [12]. Therefore, they provide a versatile tool to evaluate how the structure of the resulting complexes affects their photophysical properties.

In recent studies, our group has examined the structure–properties relationship of several Zn(II) and Cd(II) metal complexes with benzoic acid or piperonylic acid (HPip) and 3-phenylpyridine, 4-phenylpyridine and 4-acetylpyridine (4-Acpy) [13,14]. From these studies, three dimeric  $[\text{Zn}_2(\text{CO}_2)_4]$  paddle-wheel and three dimeric  $[\text{Cd}_2(\text{CO}_2)_4]-2+2$  double bridged complexes were synthesized and their crystal structures were elucidated. The UV-Vis absorption measurements identified 4-Acpy as a promising ligand to avoid ACQ by minimizing intermolecular interactions. Fluorescence experiments confirmed the minimization of PET processes through complexation by enhanced fluorescence of the complexes with respect to the free ligands. Likewise, the work performed using Hg(II), HPip and a range of pyridine derivative ligands (3-phenylpyridine, 4-phenylpyridine, 2,2'-bipyridine, 1,10-phenanthroline, 2,2':6',2'-terpyridine and di-(2-picolyl)amine) resulted in the formation of monomeric, dimeric and polymeric or oligomeric metal complexes [15]. Their UV-Vis measurements evinced that the aggregation effects can cause a significant shift on the absorption spectra, and the time-dependent density functional theory (TD-DFT) calculations displayed that the highest occupied molecular orbitals (HOMOs) were distributed along the Pip while the lowest occupied molecular orbitals (LUMOs) were over the pyridine ligands [15].

As a continuation of this work, in this paper, we have studied the reaction between  $\text{M}(\text{OAc})_2$  ( $\text{M} = \text{Zn}(\text{II})$  and  $\text{Cd}(\text{II})$ ;  $\text{OAc} = \text{acetate}$ ) with HPip and two *N*-donor ligands ( $\text{dPy} = 4\text{-acetylpyridine (4-Acpy)}$  and  $\text{isonicotinamide (Isn)}$ ), considering that 4-Acpy has a good electron withdrawing group and Isn has a better electron donor group. From these reactions we have isolated one paddle-wheel complex  $[\text{Zn}(\mu\text{-Pip})_2(4\text{-Acpy})]_2$  (1); a mixture of one dimer and two monomers  $[\text{Zn}(\mu\text{-Pip})(\text{Pip})(\text{Isn})_2]_2 \cdot 2[\text{Zn}(\text{Pip})_2(\text{HPip})(\text{Isn})] \cdot 2\text{MeOH}$  (2); and two dimeric complexes  $[\text{Cd}(\mu\text{-Pip})(\text{Pip})(4\text{-Acpy})_2]_2$  (3) and  $[\text{Cd}(\mu\text{-Pip})(\text{Pip})(\text{Isn})_2]_2 \cdot \text{MeOH}$  (4).

Besides, their molecular and supramolecular structures have been analyzed and their photophysical properties have been studied. These results have been supported by TD-DFT calculations.

## 2. Results and Discussion

### 2.1. Synthesis and General Characterization

Complexes **1–4** were prepared via a combination of  $M(\text{OAc})_2 \cdot 2\text{H}_2\text{O}$  ( $M = \text{Zn}(\text{II})$  and  $\text{Cd}(\text{II})$ ;  $\text{OAc} = \text{acetate}$ ) salts and 1,3-benzodioxole-5-carboxylic acid (Piperonylic, HPip) with  $\text{dPy} = 4\text{-acetylpyridine}$  (4-AcPy) or isonicotinamide (Isn) in MeOH as solvent at room temperature (RT). The  $M:\text{Pip}:\text{dPy}$  molar ratio used in the synthesis was 1:2:4 for complexes **1** and **3**; 1:2:1 for complex **2**; and 1:2:2 for complex **4**. The corresponding crystals suitable for X-ray crystallographic analysis were grown via slow evaporation of their mother liquors (**1**, **2** and **3**) or via recrystallization in MeOH (**4**).

All of the pure isolated metal complexes were characterized using elemental analysis, FTIR-ATR,  $^1\text{H}$ , and  $^{13}\text{C}\{^1\text{H}\}$  NMR spectroscopies, and single crystal X-ray diffraction method. In addition, their UV-Vis and fluorescence spectra were recorded and their molar absorptivity ( $\epsilon$ ) and quantum yield ( $\phi_f$ ) calculated. The elemental analysis of all of the complexes agreed with the proposed formulas.

**FTIR-ATR spectroscopy.** The FTIR-ATR spectra display the characteristic carboxylate bands of Pip ligand (S.I: Figure S1–S4). The carboxylate bands for  $\nu_{\text{as}}(\text{COO})$  appear at 1556 (**1**), 1534 and 1502 (**2**), 1539 and 1506 (**3**), or at 1539 and 1502  $\text{cm}^{-1}$  (**4**), whereas bands for  $\nu_{\text{s}}(\text{COO})$  arise at 1385 (**1**), 1385 (**2**), 1384 (**3**) and at 1385  $\text{cm}^{-1}$  (**4**). The calculated  $\Delta$  values ( $\nu_{\text{as}}(\text{COO}) - \nu_{\text{s}}(\text{COO})$ ) [16] were found to be 171 (**1**), 149 and 117 (**2**), 155 and 122 (**3**), 154 and 117 (**4**)  $\text{cm}^{-1}$ . These values suggest a bidentate bridged coordination mode of the Pip ligands ( $\mu_2\text{-}\eta^1\text{:}\eta^1$ ) in **1**, while in compounds **2–4** either bidentate bridged ( $\mu_2\text{-}\eta^1\text{:}\eta^1$ ) or bridged and chelated ( $\mu_1\text{-}\eta^2$ ) coordination modes are inferred. All of the spectra present bands attributable to the  $\nu(\text{C}=\text{O})$  belong either to the 4-AcPy at 1703  $\text{cm}^{-1}$  (**1**) and 1693  $\text{cm}^{-1}$  (**3**) or to the Isn ligand at 1701  $\text{cm}^{-1}$  (**2**) and 1704  $\text{cm}^{-1}$  (**4**). The  $\nu(\text{C}=\text{O})$  of 4-AcPy in complexes **1** and **3** is almost unvaried compared to the free 4-AcPy (1693  $\text{cm}^{-1}$ ) due to small changes in the intermolecular interactions. In contrast, the  $\nu(\text{C}=\text{O})$  of Isn in complexes **2** and **4** is blue shifted compared to the free Isn (1655  $\text{cm}^{-1}$ ), which is associated with a decrease in hydrogen bonding [17]. Besides, the  $\nu(\text{C}=\text{O})$  of the protonated HPip ligand in **2** appears at 1678  $\text{cm}^{-1}$ , slightly blue shifted compared to the free HPip (1667  $\text{cm}^{-1}$ ), suggesting that the hydrogen bonding is maintained [17]. Additional bands in all of the complexes attributable to the aromatic ring vibrations have also been identified [18,19].

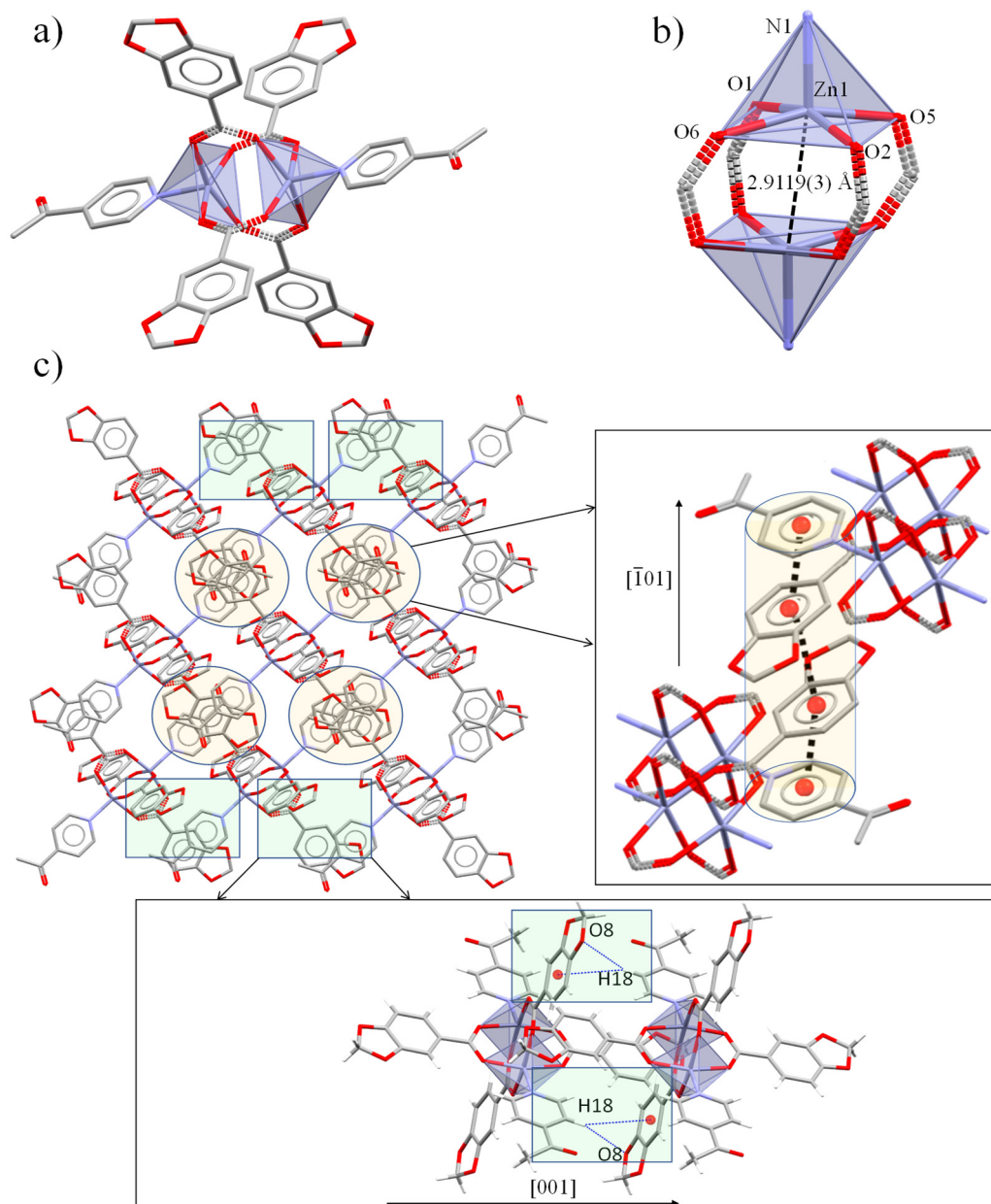
**NMR spectroscopy.** The  $^1\text{H}$  and  $^{13}\text{C}\{^1\text{H}\}$  NMR spectra were recorded in  $\text{DMSO-}d_6$ . All of the spectra show the signals belonging to Pip and either to 4-AcPy or Isn. In the  $^1\text{H}$  NMR spectra of all the compounds (S.I: Figures S5–S8), the signals attributable to the Pip ligand have been assigned to the aromatic protons between 7.55 and 6.80 ppm, and to the aliphatic protons of the dioxole unit between 6.07 and 5.95 ppm (HPip: 7.54–7.00 and 6.12 ppm). The signals from the aromatic and  $\text{CH}_3$  protons of the 4-AcPy ligand appear between 8.82 and 7.72 or between 2.63 and 2.39 ppm (4-AcPy: 8.68, 7.69 and 2.36 ppm). The aromatic signals from Isn appear between 8.71 and 7.72 ppm. The protons from the amide group appear between 8.23 and 7.66 ppm (Isn: 8.68, 8.21, 7.73, 7.69). In addition, the  $^1\text{H}$  NMR spectra of **1–4** confirms the 2:1 (**1** and **3**), 5:4 (**2**) or 1:1 (**4**) molar ratio of the Pip with regards to the 4-AcPy or Isn ligands.

In the  $^{13}\text{C}\{^1\text{H}\}$  NMR spectra, the signals corresponding to the carboxylate group of the Pip ligands have been found at 171.4 (**1**), 171.2 (**2**), 172.3 (**3**), and 172.1 (**4**) ppm (S.I: Figures S9–S12); the aromatic carbons between 150.8 and 107.5 ppm; and the aliphatic carbon of the dioxole unit between 101.8 and 101.4 ppm. The signals attributable to the  $\text{C}=\text{O}_{4\text{-AcPy}}$  at 198.1 (**1**) and 198.5 ppm (**3**), and those from the  $\text{C}=\text{O}_{\text{Isn}}$  are at 166.3 (**2**) and 166.6 ppm (**4**). Besides, those corresponding to the  $-\text{CH}_3$  group are found at 27.0 (**1**) and 27.3 ppm (**3**) [18]. The correct assignment of the C aromatic atoms required the use of DEPT-135 experiments for complexes **1** and **2**.

### 2.2. Crystal and Extended Structure of Zn(II) Complexes

$[\text{Zn}(\mu\text{-Pip})_2(4\text{-AcPy})]_2$  (**1**). It crystallizes in the triclinic P-1 space group and has a dimeric paddle-wheel structure (Figure 1a). Each Zn(II) center has a  $[\text{ZnO}_4\text{N}]$  core belonging

to the four Pip ligands and a 4-Acpy linker (Figure 1b) and exhibits a square pyramidal geometry (S.I.: Table S1) [20] with a value of  $\tau = 0.004$  [21], being pulled 0.340 Å out of the equatorial plane towards the axial position. The structure is held together by the four Pip ligands exhibiting a bidentate bridged ( $\mu_2\text{-}\eta^1\text{:}\eta^1$ ) coordination mode. The bond lengths and angles are within the range of analogous paddle-wheel complexes [13,14] (Table 1).



**Figure 1.** (a) Molecular paddle-wheel like structure of complex 1 and (b) Zn(II) metal core. (c) Supramolecular assembly forming 2D sheets promoted by  $\pi \cdots \pi$  supported by C-H  $\cdots$   $\pi$  and C-H  $\cdots$  O interactions. In detail,  $\pi \cdots \pi$  (light orange) and C-H  $\cdots$   $\pi$ /C-H  $\cdots$  O (light green) associations between Pip and 4-Acpy ligands.

**Table 1.** Bond lengths (Å), bond angles (°) and intermolecular interactions in **1**.

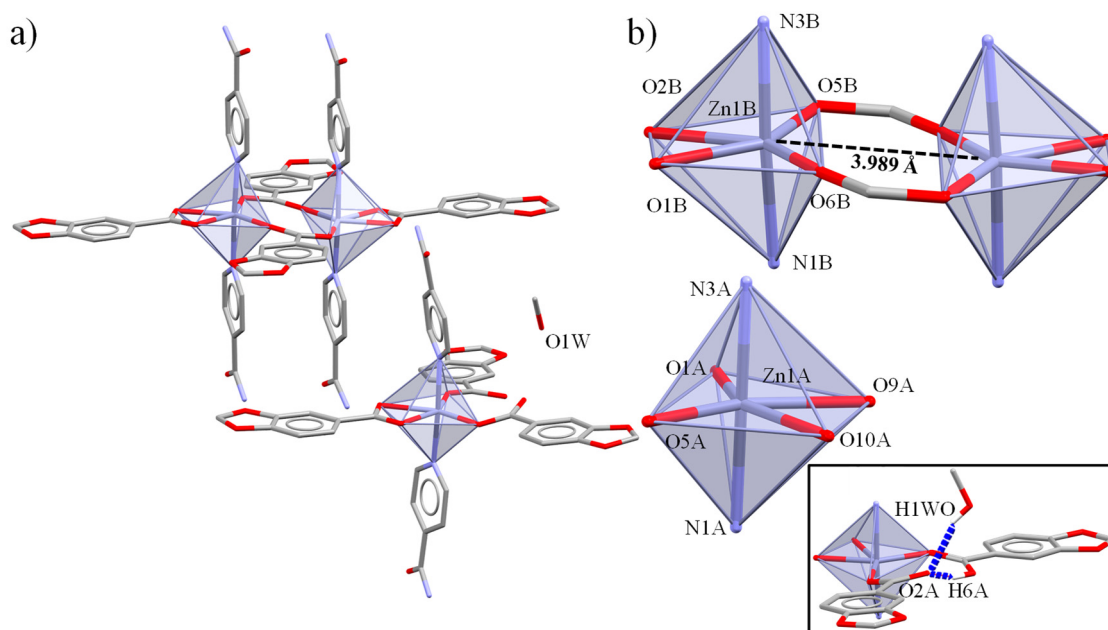
Bond Lengths					
Zn(1)-O(5)	2.0372(12)	Zn(1)-O(1)#1	2.0489(12)		
Zn(1)-O(2)	2.0388(11)	Zn(1)-N(1)	2.0338(14)		
Bond angles					
N(1)-Zn(1)-O(5)	101.93(5)	O(2)-Zn(1)-O(6)#1	87.85(5)		
N(1)-Zn(1)-O(2)	100.22(5)	N(1)-Zn(1)-O(1)#1	98.73(5)		
O(5)-Zn(1)-O(2)	89.35(5)	O(5)-Zn(1)-O(1)#1	88.98(5)		
N(1)-Zn(1)-O(6)#1	97.36(6)	O(2)-Zn(1)-O(1)#1	160.91(5)		
O(5)-Zn(1)-O(6)#1	160.70(6)	O(6)#1-Zn(1)-O(1)#1	87.47(5)		
Intermolecular interactions					
	H...A (Å)	D...A (Å)	D-H (Å)	>D-H...A (°)	
C18-H18...Cg(1)	3.395	4.273(3)	0.950	154.63	
C18-H18...O8	2.598(2)	3.467(3)	0.950	152.31	
$\pi \cdots \pi$ interactions					
Cg(I)...Cg(J)	Cg...Cg <sub>a</sub>	$\alpha$ <sup>b</sup>	$\beta, \gamma$ <sup>c</sup>	Cg(I)_Perp,Cg(J)_Perp <sup>d</sup>	Slippage <sup>e</sup>
Cg(2)...Cg(2)	3.5075(10)	0	16.5, 16.5	3.3638(7), 3.3639(7)	0.994
Cg(2)...Cg(3)	3.7861(10)	10.37(8)	31.1, 20.7	3.5410(7), 3.2430(7)	-

#1  $-x, -y+1, -z+1$ . Cg...Cg distances are given in Å. <sup>a</sup> Cg...Cg = distance between ring centroids (Å); <sup>b</sup>  $\alpha$  = dihedral angle between planes I and J (°); <sup>c</sup> Offset angles:  $\beta$  = angle Cg(I)-Cg(J) and normal to plane I (°) and  $\gamma$  = angle Cg(I)-Cg(J) and normal to plane J (°) ( $\beta = \gamma$ , when  $\alpha = 0$ ); <sup>d</sup> Perpendicular distance (Å) of Cg(I) on plane J and perpendicular distance (Å) of Cg(J) on plane I (equal when  $\alpha = 0$ ); <sup>e</sup> Slippage = Horizontal displacement or slippage between Cg(I) and Cg(J) (equal for both centroids when  $\alpha = 0$ ). Cg(1) = C10-C11-C12-C14-C15-C16; Cg(2) = C2-C3-C4-C6-C7-C8; Cg(3) = N1-C17-C18-C19-C20-C21.

These paddle-wheel units are assembled in groups of four by  $\pi \cdots \pi$  interactions between Pip (Cg(2)) and 4-Acpy (Cg(3)) ligands in a 4-Acpy-Pip-Pip-4-Acpy sequence, stacking them through the  $[\bar{1}01]$  direction. Finally, the dimers are associated in chains along the  $[1]$  direction (Figure 1c) by weak C-H... $\pi$  [22] and C-H...O interactions between the *m*-H aromatic H atom from the 4-Acpy ligand and both the aromatic ring (C18-H18...Cg(1), 3.395 Å) and O dioxole atom of Pip (C18-H18...O8, 2.598(2) Å) (Table 1, Figure 1c). All of these sets of interactions form 2D sheets of dimers along the *ac* plane.

$[\text{Zn}(\mu\text{-Pip})(\text{Pip})(\text{Isn})_2]_2 \cdot 2[\text{Zn}(\text{Pip})_2(\text{HPip})(\text{Isn})] \cdot 2\text{MeOH}$  (**2**). It crystallizes in the triclinic P-1 space group and comprises one dimeric and two monomeric structural units with two occluded MeOH molecules in the unit cell (Figure 2a). Such a structure is unusual and seems to be driven by strong interactions between the amide moieties, which also promoted the same behavior with Cu(II) [23]. Both the monomer and dimer have hexacoordinated Zn(II) atoms bearing  $[\text{ZnO}_4\text{N}_2]$  cores with a distorted octahedral geometry (S.I: Table S1) presenting an *ata* (average twist angle) of 58.41° for Zn(1A) and 58.81° for Zn(1B) [24,25] (Table 2, Figure 2b). The dimeric unit is composed of four Pip and four Isn ligands. In contrast, the monomer contains two Pip, one HPip and two Isn ligands. The Zn(II) centers in the dimer are joined by two bidentate bridged ( $\mu_2\text{-}\eta^1\text{:}\eta^1$ ) Pip ligands while the remaining two Pip units have a bidentate chelate coordination mode ( $\mu_1\text{-}\eta^2$ ). Selected bond lengths and angles are displayed in Table 2. In addition, the Pip ligands in the monomer present  $\mu_1\text{-}\eta^2$  and  $\mu_1\text{-}\eta^1$  while HPip has a  $\mu_1\text{-}\eta^1$  coordination mode. This  $\mu_1\text{-}\eta^1$  coordination of the ligands is supported by an intramolecular O-H...O between the acidic proton of HPip and the uncoordinated carboxylate oxygen atom from the neighboring ligand.





**Figure 2.** (a) Molecular structure of the dimeric and monomeric units in complex 2. (b) Representation of the dimeric and monomeric Zn (II) cores. Inset of the intra- and intermolecular interactions between HPip and MeOH with  $\mu_1\text{-}\eta^1\text{-Pip}$  ligand.

**Table 2.** Bond lengths (Å), bond angles (°) and intramolecular interactions in 2.

Monomer			
Bond lengths			
Zn(1A)-O(1A)	2.001(3)	Zn(1A)-N(3A)	2.154(3)
Zn(1A)-O(5A)	2.052(3)	Zn(1A)-O(9A)	2.184(3)
Zn(1A)-N(1A)	2.140(3)	Zn(1A)-O(10A)	2.257(3)
Bond angles			
O(1A)-Zn(1A)-O(5A)	117.34(11)	N(1A)-Zn(1A)-O(9A)	91.03(11)
O(1A)-Zn(1A)-N(1A)	91.70(12)	N(3A)-Zn(1A)-O(9A)	89.91(10)
O(5A)-Zn(1A)-N(1A)	88.62(11)	O(1A)-Zn(1A)-O(10A)	151.96(10)
O(1A)-Zn(1A)-N(3A)	92.36(11)	O(5A)-Zn(1A)-O(10A)	90.69(10)
O(5A)-Zn(1A)-N(3A)	88.51(11)	N(1A)-Zn(1A)-O(10A)	88.07(11)
N(1A)-Zn(1A)-N(3A)	175.79(12)	N(3A)-Zn(1A)-O(10A)	88.89(10)
O(1A)-Zn(1A)-O(9A)	92.89(10)	O(9A)-Zn(1A)-O(10A)	59.09(9)
O(5A)-Zn(1A)-O(9A)	149.77(11)		
TA value (°)		ATA value (°)	
O(10A)-Cg(3)-Cg(4)-N(1A)	46.72		
N(3A)-Cg(3)-Cg(4)-O(5A)	64.48		58.41
O(9A)-Cg(3)-Cg(4)-O(1A)	64.04		

Table 2. Cont.

Monomer					
Dimer					
Bond lengths					
Zn(1B)-O(5B)	2.025(3)	Zn(1B)-N(1B)	2.153(3)		
Zn(1B)-O(6B)#1	2.034(3)	Zn(1B)-O(2B)	2.271(3)		
Zn(1B)-O(1B)	2.140(3)	Zn(1B)···Zn(1B)	3.9893(6)		
Zn(1B)-N(3B)	2.151(3)				
Bond angles					
O(5B)-Zn(1B)-O(6B)#1	125.03(11)	O(1B)-Zn(1B)-N(1B)	93.02(11)		
O(5B)-Zn(1B)-O(1B)	145.44(11)	N(3B)-Zn(1B)-N(1B)	176.78(12)		
O(6B)#1-Zn(1B)-O(1B)	89.46(11)	O(5B)-Zn(1B)-O(2B)	86.09(10)		
O(5B)-Zn(1B)-N(3B)	88.48(11)	O(6B)#1-Zn(1B)-O(2B)	148.82(11)		
O(6B)#1-Zn(1B)-N(3B)	89.27(12)	O(1B)-Zn(1B)-O(2B)	59.55(10)		
O(1B)-Zn(1B)-N(3B)	90.17(11)	N(3B)-Zn(1B)-O(2B)	93.82(11)		
O(5B)-Zn(1B)-N(1B)	89.12(11)	N(1B)-Zn(1B)-O(2B)	88.17(11)		
O(6B)#1-Zn(1B)-N(1B)	90.34(12)				
Intramolecular interactions					
	H···A (Å)	D···A (Å)	D-H (Å)	>D-H···A (°)	
O(6A)-H(6A)···O(2A)	1.723	2.503(4)	0.840	153.7	
$\pi$ ··· $\pi$ interactions					
Cg(I)···Cg(J)	Cg···Cg <sup>a</sup>	$\alpha$ <sup>b</sup>	$\beta$ , $\gamma$ <sup>c</sup>	Cg(I)_Perp, Cg(J)_Perp <sup>d</sup>	Slippage <sup>e</sup>
Cg(1)···Cg(2)	3.884(2)	1.7(2)	15.1, 16.1	3.7307(17), 3.7493(17)	-
TA value (°)			ATA value (°)		
N(3B)-Cg(5)-Cg(6)-O(5B)		68.38			
O(1B)-Cg(5)-Cg(6)-O(6B)		62.50			
O(2B)-Cg(5)-Cg(6)-N(1B)		45.55			

#1  $-x+1, -y+1, -z$ . Cg···Cg distances are given in Å. <sup>a</sup> Cg···Cg = distance between ring centroids (Å); <sup>b</sup>  $\alpha$  = dihedral angle between planes I and J (°); <sup>c</sup> Offset angles:  $\beta$  = angle Cg(I)-Cg(J) and normal to plane I (°) and  $\gamma$  = angle Cg(I)-Cg(J) and normal to plane J (°) ( $\beta = \gamma$ , when  $\alpha = 0$ ); <sup>d</sup> Perpendicular distance (Å) of Cg(I) on plane J and perpendicular distance (Å) of Cg(J) on plane I (equal when  $\alpha = 0$ ); <sup>e</sup> Slippage = Horizontal displacement or slippage between Cg(I) and Cg(J) (equal for both centroids when  $\alpha = 0$ ). Cg(1) = N1B-C17B-C18B-C19B-C20B-C21B; Cg(2) = N3B-C23B-C24B-C25B-C26B-C27B. Cg(3) = N3A-O9A-O10A; Cg(4) = N1A-O1A-O5A; Cg(5) = N3B-O2B-O1B; Cg(6) = N1B-O6B-O5B.

Simultaneously, the carboxylate O atom interacts with the MeOH molecule through an intermolecular O-H···O hydrogen bond. Besides, the aromatic rings from Isn stacks in an intramolecular  $\pi$ ··· $\pi$  interaction (Table 2).

The monomers and dimers form a 3D supramolecular net via amide···amide, N-H/C-H···O and  $\pi$ ··· $\pi$  interactions (Table 3). Amide···amide piles monomers and dimers between themselves while Isn, through its N-H<sub>anti</sub> and *m*-H, form a dimer via association with an O atom from a chelate Pip ligand of a monomer and vice versa along [22] (Figure 3a). In addition,  $\pi$ ··· $\pi$  interactions between Pip ligands stack monomers and dimers in an ordered sequence from Cg(3) to Cg(7) along the *b* axis (Figure 3b).

**Table 3.** Intermolecular interactions in **2**.

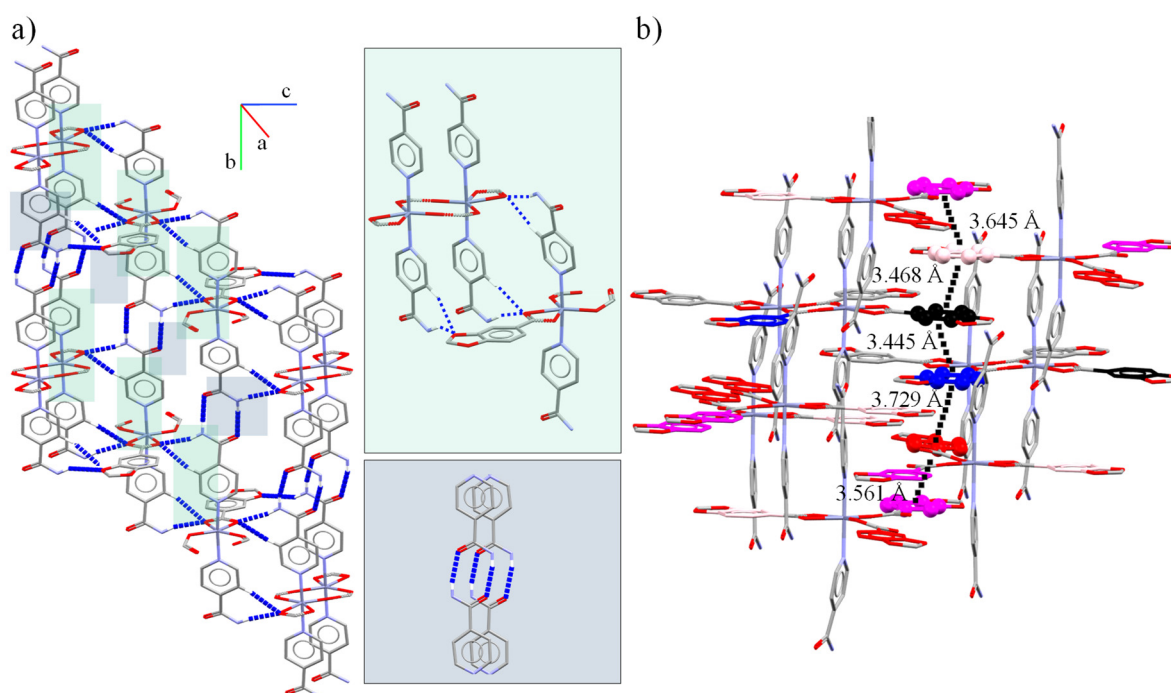
Intermolecular Interactions.	H···A (Å)	D···A (Å)	D-H (Å)	>D-H···A (°)	
N(2A)-H(2AA)···O(14A)	2.07	2.934(4)	0.88	168.4	
N(2A)-H(2BA)···O(9A)	2.05	2.903(4)	0.88	164	
N(4A)-H(4AA)···O(13A)	2.04	2.921(4)	0.88	176.1	
N(4A)-H(4BA)···O(2B)	2.02	2.855(4)	0.88	157	
N(2B)-H(2BA)···O(10B)	1.99	2.858(4)	0.88	166.8	
N(2B)-H(2BB)···O(12A)	2.46	3.268(5)	0.88	153.6	
N(4B)-H(4BA)···O(9B)	2.05	2.913(4)	0.88	165.9	
N(4B)-H(4BB)···O(10A)	2.01	2.868(4)	0.88	164.7	
O(1W)-H(1WO)···O(2A)	2.10	2.900(6)	0.84	160.1	
$\pi \cdots \pi$ interactions					
Cg(I)···Cg(J) Cg···Cg <sup>a</sup>	$\alpha$ <sup>b</sup>	$\beta, \gamma$ <sup>c</sup>	Cg(I)_Perp, Cg(J)_Perp <sup>d</sup>	Slippage <sup>e</sup>	
Cg(7)···Cg(6)	3.445(2)	1.24(18)	10.2, 9.0	3.4022(16), 3.3905(14)	-
Cg(7)···Cg(3)	3.468(2)	3.01(18)	13.6, 15.4	3.3432(16), 3.3710(14)	-
Cg(6)···Cg(5)	3.729(2)	4.64(17)	25.5, 22.5	3.4456(14), 3.3670(15)	-
Cg(5)···Cg(4)	3.561(2)	2.91(18)	16.2, 18.6	3.3740(15), 3.4202(16)	-
Cg(4)···Cg(3)	3.645(2)	2.60(18)	24.2, 21.7	3.3852(16), 3.3244(14)	-

Cg···Cg distances are given in Å. <sup>a</sup> Cg···Cg = distance between ring centroids (Å); <sup>b</sup>  $\alpha$  = dihedral angle between Planes I and J (°); <sup>c</sup> Offset angles:  $\beta$  = angle Cg(I)-Cg(J) and normal to plane I (°) and  $\gamma$  = angle Cg(I)-Cg(J) and normal to plane J (°) ( $\beta = \gamma$ , when  $\alpha = 0$ ); <sup>d</sup> Perpendicular distance (Å) of Cg(I) on plane J and perpendicular distance (Å) of Cg(J) on plane I (equal when  $\alpha = 0$ ); <sup>e</sup> Slippage = Horizontal displacement or slippage between Cg(I) and Cg(J) (equal for both centroids when  $\alpha = 0$ ). Cg(3) = C18A-C19A-C20A-C21A-C23A-C24A; Cg(4) = C10A-C11A-C12A-C13A-C15A-C16A; Cg(5) = C2A-C3A-C4A-C5A-C7A-C8A; Cg(6) = C10B-C11B-C12B-C13B-C15B-C16B; Cg(7) = C2B-C3B-C4B-C5B-C7B-C8B.

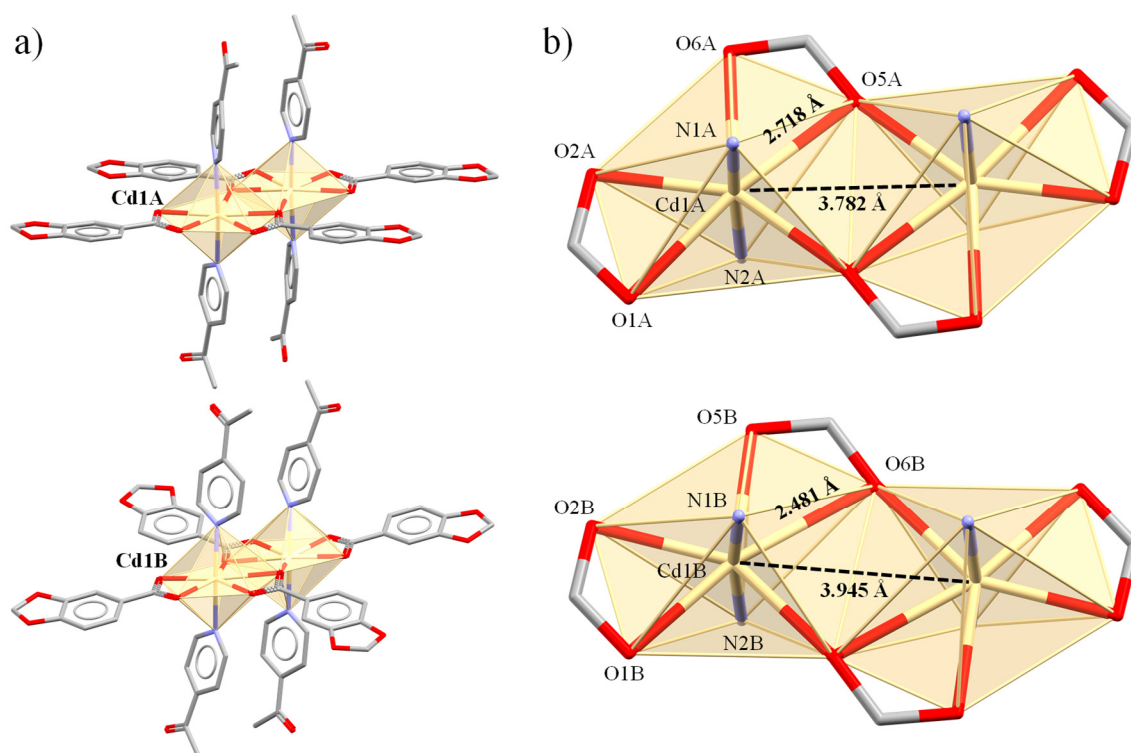
### 2.3. Crystal and Extended Structure of Cd(II) Compounds

[Cd( $\mu$ -Pip)(Pip)(4-AcPy)<sub>2</sub>]<sub>2</sub> (**3**) and [Cd( $\mu$ -Pip)(Pip)(Isn)<sub>2</sub>]<sub>2</sub>·MeOH (**4**). They crystallize in the triclinic P-1 and monoclinic P2/c space groups, respectively, and contain two crystallographically independent molecules in the unit cell (molecules A and B). Both complexes have a dimeric structure with four Pip ligands and four 4-AcPy (**3**) (Figure 4a,b) or four Isn (**4**) (Figure 5a,b), in which the heptacoordinated Cd atoms present a [CdO<sub>5</sub>N<sub>2</sub>] core with a distorted pentagonal-bipyramidal geometry (S.I.: Table S1). The bond lengths and angles are within the range of analogous dimeric Cd(II) complexes [13,14] (Tables 4 and 5).

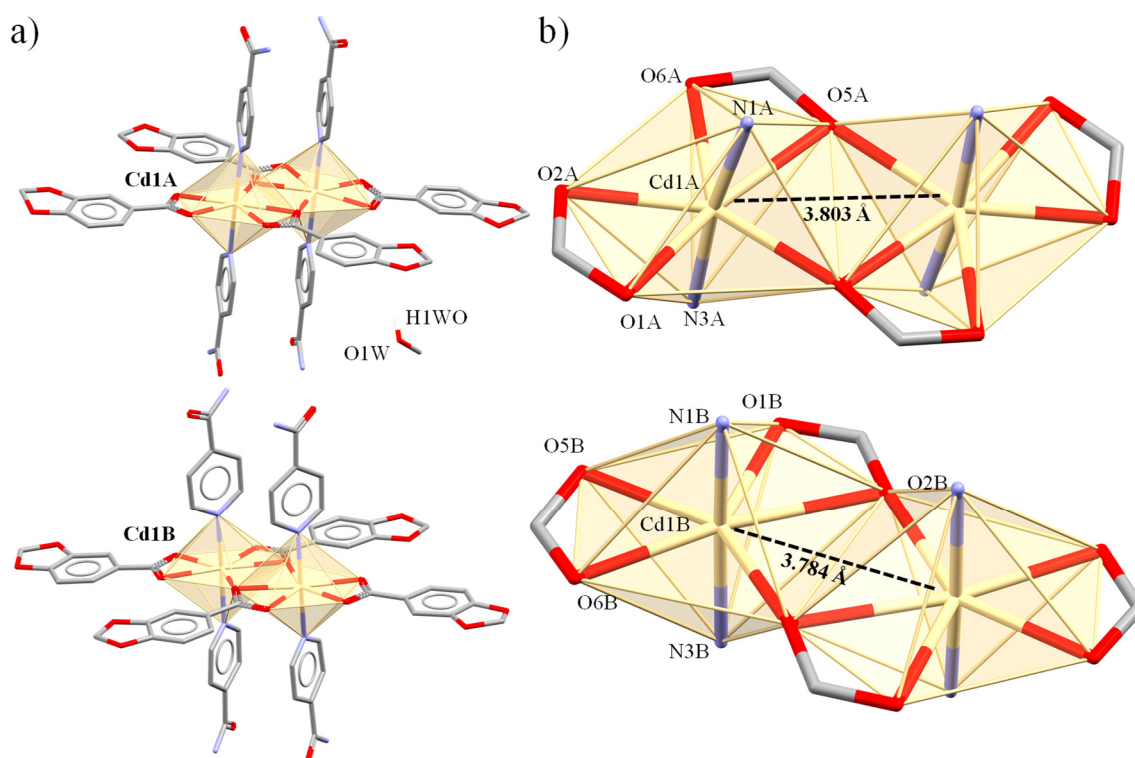




**Figure 3.** Supramolecular interactions in **2** (a) promoted by Isn, Pip and HPip ligands. In detail, reciprocal N-H $\cdots$ O and N-H<sub>anti</sub>/C-H $\cdots$ O interactions. (b)  $\pi \cdots \pi$  interactions between Pip aromatic rings highlighted in violet (Cg(4)), pink (Cg(3)), black (Cg(7)), blue (Cg(6)), and red (Cg(5)).



**Figure 4.** Molecular structure of the crystallographically independent dimeric units in complex **3**. Units A and B (a) with their corresponding Cd(II) cores (b).



**Figure 5.** Molecular structure of the crystallographically independent dimeric units in complex **4**. Units A and B (a) with their corresponding Cd(II) cores (b).

**Table 4.** Bond lengths (Å), bond angles (°) and intramolecular  $\pi \cdots \pi$  interactions in **3**.

Molecule A			
Bond lengths			
Cd(1A)-O(2A)	2.285(2)	Cd(1A)-N(2A)	2.322(3)
Cd(1A)-O(5A)	2.330(2)	Cd(1A)-O(6A)#1	2.395(2)
Cd(1A)-O(5A)#1	2.482(2)	Cd(1A)-O(1A)	2.485(2)
Cd(1A)-N(1A)	2.332(3)	Cd(1A) $\cdots$ Cd(1A)	3.7820(5)
Bond angles			
O(2A)-Cd(1A)-N(2A)	93.46(9)	N(2A)-Cd(1A)-O(5A)#1	92.29(8)
O(2A)-Cd(1A)-O(5A)	149.25(7)	O(5A)-Cd(1A)-O(5A)#1	76.42(8)
N(2A)-Cd(1A)-O(5A)	87.09(8)	N(1A)-Cd(1A)-O(5A)#1	90.58(8)
O(2A)-Cd(1A)-N(1A)	89.75(9)	O(6A)#1-Cd(1A)-O(5A)#1	53.41(7)
N(2A)-Cd(1A)-N(1A)	172.17(9)	O(2A)-Cd(1A)-O(1A)	54.75(7)
O(5A)-Cd(1A)-N(1A)	86.51(8)	N(2A)-Cd(1A)-O(1A)	87.86(8)
O(2A)-Cd(1A)-O(6A)#1	80.79(7)	O(5A)-Cd(1A)-O(1A)	94.61(7)
N(2A)-Cd(1A)-O(6A)#1	95.65(8)	N(1A)-Cd(1A)-O(1A)	88.17(8)
O(5A)-Cd(1A)-O(6A)#1	129.79(7)	O(6A)#1-Cd(1A)-O(1A)	135.54(7)
O(2A)-Cd(1A)-O(5A)#1	134.18(7)	O(5A)#1-Cd(1A)-O(1A)	171.00(7)

Table 4. Cont.

Molecule B					
Bond lengths					
Cd(1B)-O(5B)	2.265(2)	Cd(1B)-N(1B)	2.318(3)		
Cd(1B)-O(6B)#2	2.317(2)	Cd(1B)-N(2B)	2.335(3)		
Cd(1B)-O(1B)	2.366(2)	Cd(1B)-O(2B)	2.367(2)		
Cd(1B)-O(6B)	2.718(2)	Cd(1B)···Cd(1B)	3.9452(6)		
Bond angles					
O(5B)-Cd(1B)-O(6B)#2	127.84(8)	N(1B)-Cd(1B)-O(1B)	93.92(8)		
O(5B)-Cd(1B)-N(1B)	96.03(9)	N(2B)-Cd(1B)-O(1B)	86.04(8)		
O(6B)#2-Cd(1B)-N(1B)	85.79(8)	O(5B)-Cd(1B)-O(2B)	85.45(8)		
O(5B)-Cd(1B)-N(2B)	88.90(8)	O(6B)#2-Cd(1B)-O(2B)	145.20(8)		
O(6B)#2-Cd(1B)-N(2B)	86.31(8)	N(1B)-Cd(1B)-O(2B)	102.18(9)		
N(1B)-Cd(1B)-N(2B)	172.09(9)	N(2B)-Cd(1B)-O(2B)	84.33(9)		
O(5B)-Cd(1B)-O(1B)	141.05(7)	O(1B)-Cd(1B)-O(2B)	55.62(7)		
O(6B)#2-Cd(1B)-O(1B)	90.36(8)				
$\pi \cdots \pi$ interactions					
Cg(I)···Cg(J)	Cg···Cg <sup>a</sup>	$\alpha$ <sup>b</sup>	$\beta, \gamma$ <sup>c</sup>	Cg(I)_Perp, Cg(J)_Perp <sup>d</sup>	Slippage <sup>e</sup>
Cg(1)···Cg(2)	3.7430	16.0	25.0, 10.4	3.6813, 3.3912	-
Cg(3)···Cg(4)	3.8415	13.0	21.2, 18.6	3.6418, 3.5810	-

#1  $-x, -y+2, -z, \#2 -x+1, -y, -z+1$ . Cg···Cg distances are given in Å. <sup>a</sup> Cg···Cg = distance between ring centroids (Å); <sup>b</sup>  $\alpha$  = dihedral angle between planes I and J (°); <sup>c</sup> Offset angles:  $\beta$  = angle Cg(I)-Cg(J) and normal to plane I (°) and  $\gamma$  = angle Cg(I)-Cg(J) and normal to plane J (°) ( $\beta = \gamma$ , when  $\alpha = 0$ ); <sup>d</sup> Perpendicular distance (Å) of Cg(I) on plane J and perpendicular distance (Å) of Cg(J) on plane I (equal when  $\alpha = 0$ ); <sup>e</sup> Slippage = Horizontal displacement or slippage between Cg(I) and Cg(J) (equal for both centroids when  $\alpha = 0$ ). Cg(1) = N1A-C17A-C18A-C19A-C20A-C21A; Cg(2) = N2A-C24A-C25A-C26A-C27A-C28A; Cg(3) = N2B-C24B-C25B-C26B-C27B-C28B; Cg(4) = N1B-C17B-C18B-C19B-C20B-C21B.

Table 5. Bond lengths (Å), bond angles (°) and intramolecular  $\pi \cdots \pi$  interactions of 4.

Molecule A					
Bond lengths					
Cd(1A)-N(3A)	2.307(3)	Cd(1A)-O(6A)#1	2.360(2)		
Cd(1A)-N(1A)	2.331(3)	Cd(1A)-O(1A)	2.427(3)		
Cd(1A)-O(5A)	2.337(2)	Cd(1A)-O(5A)#1	2.524(2)		
Cd(1A)-O(2A)	2.354(2)	Cd(1A)···Cd(1A)	3.8026(6)		
Bond angles					
N(3A)-Cd(1A)-N(1A)	173.41(9)	N(1A)-Cd(1A)-O(1A)	85.03(9)		
N(3A)-Cd(1A)-O(5A)	87.98(8)	O(5A)-Cd(1A)-O(1A)	92.34(8)		
N(1A)-Cd(1A)-O(5A)	85.46(8)	O(2A)-Cd(1A)-O(1A)	54.32(8)		
N(3A)-Cd(1A)-O(2A)	93.79(8)	O(6A)#1-Cd(1A)-O(1A)	136.98(8)		
N(1A)-Cd(1A)-O(2A)	91.36(9)	N(3A)-Cd(1A)-O(5A)#1	90.99(8)		
O(5A)-Cd(1A)-O(2A)	146.66(8)	N(1A)-Cd(1A)-O(5A)#1	88.15(8)		
N(3A)-Cd(1A)-O(6A)#1	91.19(8)	O(5A)-Cd(1A)-O(5A)#1	77.08(8)		
N(1A)-Cd(1A)-O(6A)#1	93.53(8)	O(2A)-Cd(1A)-O(5A)#1	136.07(7)		
O(5A)-Cd(1A)-O(6A)#1	130.48(8)	O(6A)#1-Cd(1A)-O(5A)#1	53.42(7)		

Table 5. Cont.

O(2A)-Cd(1A)-O(6A)#1	82.81(8)	O(1A)-Cd(1A)-O(5A)#1	167.86(8)		
N(3A)-Cd(1A)-O(1A)	94.64(9)				
<b>Molecule B</b>					
Bond lengths					
Cd(1B)-O(2B)#2	2.306(2)	Cd(1B)-O(5B)	2.362(2)		
Cd(1B)-N(3B)	2.321(2)	Cd(1B)-O(6B)	2.411(2)		
Cd(1B)-O(1B)	2.328(2)	Cd(1B)-O(2B)	2.572(2)		
Cd(1B)-N(1B)	2.329(3)	Cd(1B)···Cd(1B)	3.7838(6)		
Bond angles					
O(2B)#2-Cd(1B)-N(3B)	86.45(8)	N(3B)-Cd(1B)-O(6B)	86.89(8)		
O(2B)#2-Cd(1B)-O(1B)	131.21(8)	O(1B)-Cd(1B)-O(6B)	137.68(8)		
N(3B)-Cd(1B)-O(1B)	92.87(9)	N(1B)-Cd(1B)-O(6B)	91.31(8)		
O(2B)#2-Cd(1B)-N(1B)	86.99(8)	O(5B)-Cd(1B)-O(6B)	54.62(7)		
N(3B)-Cd(1B)-N(1B)	173.17(9)	O(2B)#2-Cd(1B)-O(2B)	78.31(8)		
O(1B)-Cd(1B)-N(1B)	92.89(9)	N(3B)-Cd(1B)-O(2B)	91.92(8)		
O(2B)#2-Cd(1B)-O(5B)	145.66(8)	O(1B)-Cd(1B)-O(2B)	52.93(8)		
N(3B)-Cd(1B)-O(5B)	91.89(8)	N(1B)-Cd(1B)-O(2B)	88.63(8)		
O(1B)-Cd(1B)-O(5B)	83.13(8)	O(5B)-Cd(1B)-O(2B)	136.03(7)		
N(1B)-Cd(1B)-O(5B)	92.44(8)	O(6B)-Cd(1B)-O(2B)	169.35(7)		
O(2B)#2-Cd(1B)-O(6B)	91.05(8)				
$\pi \cdots \pi$ interactions					
Cg(I)···Cg(J)	Cg···Cg <sup>a</sup>	$\alpha$ <sup>b</sup>	$\beta, \gamma$ <sup>c</sup>	Cg(I)_Perp, Cg(J)_Perp <sup>d</sup>	Slippage <sup>e</sup>
Cg(1)···Cg(1)	3.6090(19)	1	14.9	3.4870(14)	-
Cg(2)···Cg(2)	3.6435(18)	0	18.3	3.4597(13)	1.142
Cg(3)···Cg(3)	3.7740(19)	1	24.4	3.4372(14)	-
Cg(4)···Cg(4)	3.606(2)	3	2.5	3.6021(16), 3.6023(16)	-

#1  $-x + 1, y, -z + 1/2$ , #2  $-x, y, -z + 1/2$ . Cg···Cg distances are given in Å. <sup>a</sup> Cg···Cg = distance between ring centroids (Å); <sup>b</sup>  $\alpha$  = dihedral angle between Planes I and J (°); <sup>c</sup> Offset angles:  $\beta$  = angle Cg(I)-Cg(J) and normal to plane I (°) and  $\gamma$  = angle Cg(I)-Cg(J) and normal to plane J (°) ( $\beta = \gamma$ , when  $\alpha = 0$ ); <sup>d</sup> Perpendicular distance (Å) of Cg(I) on plane J and perpendicular distance (Å) of Cg(J) on plane I (equal when  $\alpha = 0$ ); <sup>e</sup> Slippage = Horizontal displacement or slippage between Cg(I) and Cg(J) (equal for both centroids when  $\alpha = 0$ ). Cg(1) = N1A-C17A-C18A-C19A-C20A-C21A; Cg(2) = N3A-C23A-C24A-C25A-C26A-C27A; Cg(3) = N1B-C17B-C18B-C19B-C20B-C21B; Cg(4) = N3B-C23B-C24B-C25B-C26B-C27B.

These dimeric arrays are held together by the two Pip ligands, displaying both bidentate bridge and chelate coordination modes ( $\mu_2\text{-}\eta^2\text{:}\eta^1$ ), and the remaining two Pip linkers complete the coordination sphere by a bidentate chelate ( $\mu_1\text{-}\eta^2$ ) coordination mode. Both dimers (**3** and **4**) present intramolecular  $\pi \cdots \pi$  interactions between the aromatic rings of 4-Acpy and Isn, respectively (Tables 6 and 7).

**Table 6.** Intermolecular interactions in 3.

Intermolecular Interactions	H···A (Å)	D···A (Å)	D-H (Å)	>D-H···A (°)
C6B-H6BB···O10A	2.45	3.24(1)	0.99	135.9
C14B-H14B···O10A	2.497	3.06(1)	0.99	115.7
C25A-H25A···O2A	2.241	3.190(5)	0.95	177.2
C27B-H27B···O2B	2.365	3.228(5)	0.95	150.8
C30A-H30A···O6A	2.371	3.317(5)	0.98	161.9
C30B-H30E···O5B	2.480	3.423(6)	0.98	161.3
$\pi \cdots \pi$ interactions				
Cg(I)···Cg(J) Cg···Cg <sup>a</sup>	$\alpha$ <sup>b</sup>	$\beta, \gamma$ <sup>c</sup>	Cg(I)_Perp, Cg(J)_Perp <sup>d</sup>	Slippage <sup>e</sup>
Cg(1)···Cg(7) 3.8236	9.0	25.6, 17.1	3.6555, 3.4478	-
Cg(5)···Cg(6) 3.7876	1.0	29.0, 28.4	3.3325, 3.3129	-
Cg(6)···Cg(6) 3.5058	0	18.3, 18.3	3.3287, 3.3287	1.100

Cg···Cg distances are given in Å. <sup>a</sup> Cg···Cg = distance between ring centroids (Å); <sup>b</sup>  $\alpha$  = dihedral angle between Planes I and J (°); <sup>c</sup> Offset angles:  $\beta$  = angle Cg(I)-Cg(J) and normal to plane I (°) and  $\gamma$  = angle Cg(I)-Cg(J) and normal to plane J (°) ( $\beta = \gamma$ , when  $\alpha = 0$ ); <sup>d</sup> Perpendicular distance (Å) of Cg(I) on plane J and perpendicular distance (Å) of Cg(J) on plane I (equal when  $\alpha = 0$ ); <sup>e</sup> Slippage = Horizontal displacement or slippage between Cg(I) and Cg(J) (equal for both centroids when  $\alpha = 0$ ). Cg(1) = N1A-C17A-C18A-C19A-C20A-C21A; Cg(5) = C2A-C3A-C4A-C5A-C7A-C8A; Cg(6) = C10A-C11A-C12A-C14A-C15A-C16A; Cg(7) = C2B-C3B-C4B-C5B-C7B-C8B.

**Table 7.** Intermolecular interactions in 4.

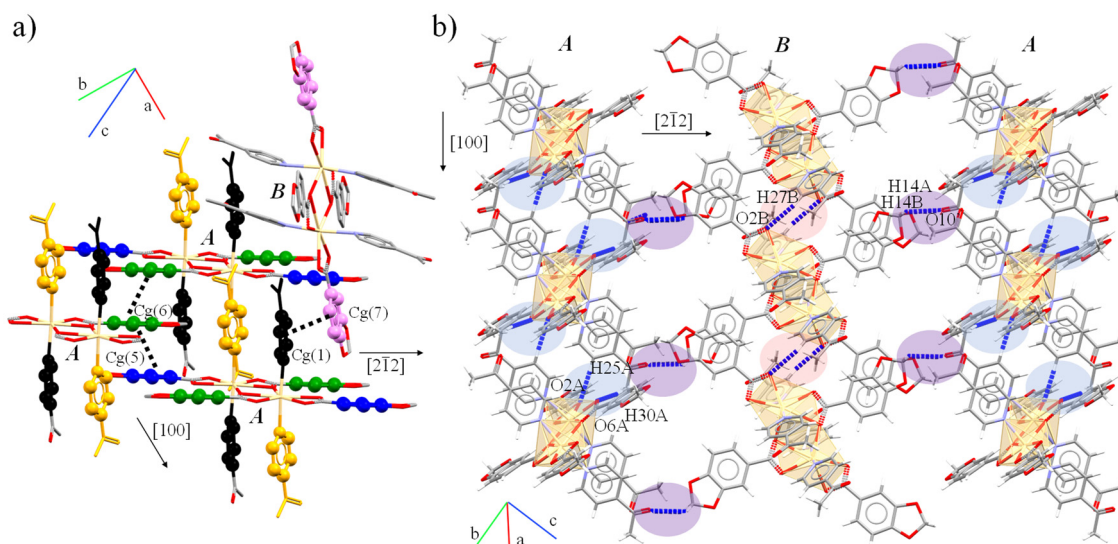
Intermolecular Interactions.	H···A (Å)	D···A (Å)	D-H (Å)	>D-H···A (°)
N2A-H2AA···O10A	1.997	2.873(3)	0.88	174.1
N2A-H2AB···O4A	2.161	2.909(4)	0.88	142.6
N4A-H4AA···O9A	2.046	2.899(3)	0.88	163.1
N4A-H4AB···O5B	2.015	2.849(3)	0.88	157.8
N4B-H4BA···O9B	2.050	2.928(4)	0.88	175.7
N4B-H4BB···O1W	2.172	2.980(8)	0.88	152.7
N2B-H2BA···O10B	2.011	2.874(4)	0.88	166.2
N2B-H2BB···O2A	2.022	2.832(3)	0.88	152.5
C26A-H26A···O5B	2.418	3.315(4)	0.95	157.3
$\pi \cdots \pi$ interactions				
Cg(I)···Cg(J) Cg···Cg <sup>a</sup>	$\alpha$ <sup>b</sup>	$\beta, \gamma$ <sup>c</sup>	Cg(I)_Perp, Cg(J)_Perp <sup>d</sup>	Slippage <sup>e</sup>
Cg(8)···Cg(7) 3.560(2)	6.19(18)	8.2, 13.4	3.4625(14), 3.5233(17)	-
Cg(5)···Cg(6) 3.824(2)	2.39(16)	26.2, 27.7	3.3857(14), 3.4303(13)	-
Cg(6)···Cg(7) 3.534(2)	1.95(17)	17.3, 15.6	3.4046(13), 3.3746(17)	-

Cg···Cg distances are given in Å. <sup>a</sup> Cg···Cg = distance between ring centroids (Å); <sup>b</sup>  $\alpha$  = dihedral angle between Planes I and J (°); <sup>c</sup> Offset angles:  $\beta$  = angle Cg(I)-Cg(J) and normal to plane I (°) and  $\gamma$  = angle Cg(I)-Cg(J) and normal to plane J (°) ( $\beta = \gamma$ , when  $\alpha = 0$ ); <sup>d</sup> Perpendicular distance (Å) of Cg(I) on plane J and perpendicular distance (Å) of Cg(J) on plane I (equal when  $\alpha = 0$ ); <sup>e</sup> Slippage = Horizontal displacement or slippage between Cg(I) and Cg(J) (equal for both centroids when  $\alpha = 0$ ). Cg(5) = C10A-C11A-C12A-C14A-C15A-C16A; Cg(6) = C2B-C3B-C4B-C5B-C7B-C8B; Cg(7) = C10B-C11B-C12B-C14B-C15B-C16B; Cg(8) = C2A-C3A-C4A-C5A-C7A-C8A.

The intermolecular assembly of **3** is guided by  $\pi \cdots \pi$  interactions supported by three pairs of weak C-H···O associations between the Pip and 4-Acyp ligands. The A dimers are stacked via planar interactions in trios, complemented by a reciprocal C-H···O between



*m*-H and H<sub>methyl</sub> from 4-AcPy and two carboxylate O atoms forming chains along the [100] direction. In contrast, B dimers only assemble via the equivalent C-H···O between Pip and 4-AcPy in the same axis. Finally, A···B dimers interact in pairs through  $\pi$ ··· $\pi$  interactions between 4-AcPy and Pip ligands supported by C-H···O between the methylene H atoms from Pip and the O atoms from 4-AcPy along [2 $\bar{1}$ 2]. Both C-H···O and  $\pi$ ··· $\pi$  interactions form 2D layers in (102) (Figure 6a,b).



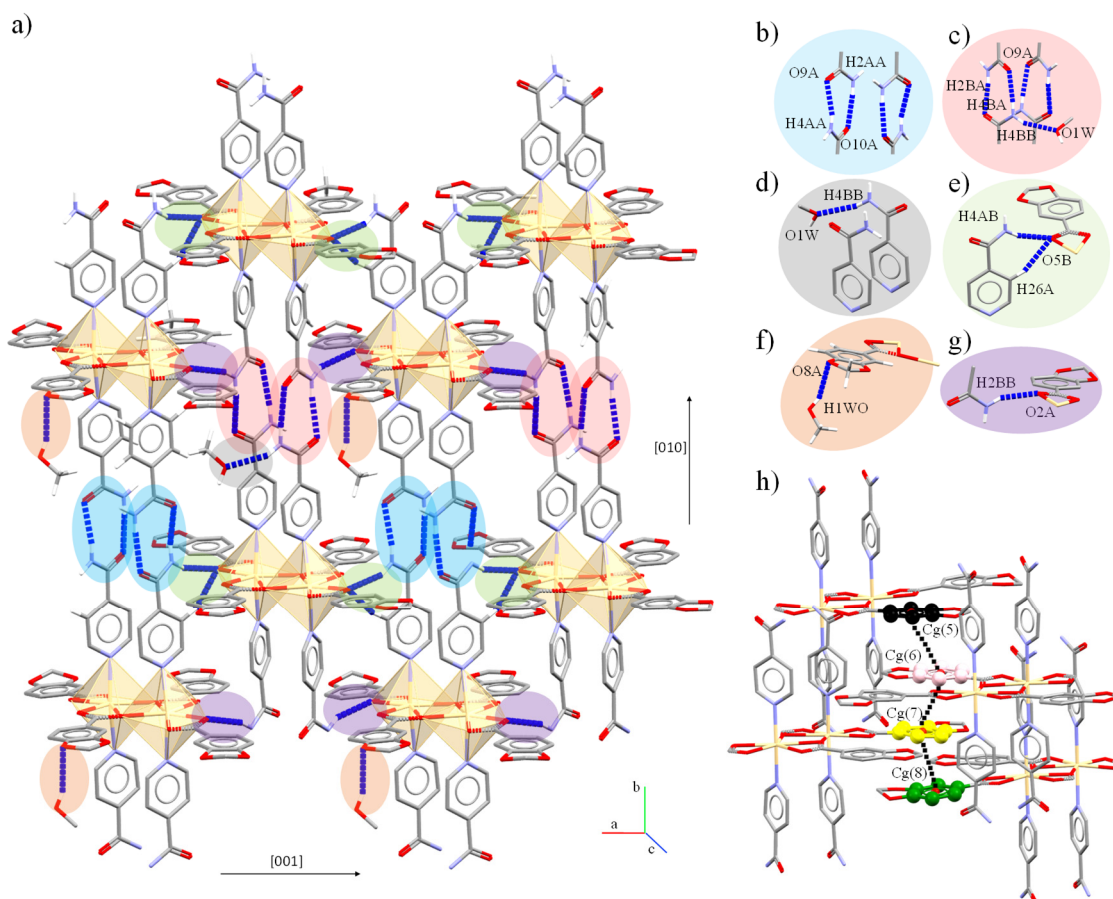
**Figure 6.** Intermolecular (a)  $\pi$ ··· $\pi$  and (b) C-H···O interactions between A and B dimers present in **3**. Color codes: Cg(1) black; Cg(2) orange, Cg(5) dark blue; Cg(6) dark green; Cg(7) violet.

The dimeric units in **4** are held together by the amide···amide pattern engaged in a head-to-head disposition which orders the dimers in chains (A···A and B···B) parallel to the [10] direction. Besides, the N-H<sub>anti</sub> also partake in a double N-H···O and *m*-C-H···O associating the chains between A···B dimers (Figure 7a–g). Planar  $\pi$ ··· $\pi$  interactions support the assembly of dimers along [10] direction (Figure 7h).

#### 2.4. Structure and Geometric Evaluation

Detailed analysis of the geometric distortions present in Zn(II) and Cd(II) complexes have been performed using SHAPE [20,26] through the *S* parameter (S.I: Table S1). A recent structural search in the CCDC [27] has revealed that Zn(II) and Cd(II) mainly present coordination numbers ranging from 4 to 7, being scarce 3 and 8, the latter only affordable by Cd(II). The predominant structural motifs reported for Zn(II) are the dimeric  $\mu$ -bridged [Zn<sub>2</sub>(CO<sub>2</sub>)<sub>2</sub>] units and the dimeric paddle-wheel [Zn<sub>2</sub>(CO<sub>2</sub>)<sub>4</sub>], bearing tetrahedral and octahedral geometries in [Zn<sub>2</sub>(CO<sub>2</sub>)<sub>2</sub>], or square pyramidal environments in [Zn<sub>2</sub>(CO<sub>2</sub>)<sub>4</sub>]. In turn, the most common motif in Cd(II) is the double bridged [Cd<sub>2</sub>(CO<sub>2</sub>)<sub>2</sub>], completed by two chelate ligands to form the [Cd<sub>2</sub>(CO<sub>2</sub>)<sub>2</sub>]-2+2 core and bearing coordination number 7 [12]. From these data it can be expected that complex **1** exhibits an almost ideal square pyramidal geometry (SPY-5, *S* = 0.240) [20] supported by the paddle-wheel structure that minimizes steric repulsion between the ligands. In the case of **2**, both the monomer and dimer present uncommon structural motifs as the monomeric [Zn(CO<sub>2</sub>)<sub>3</sub>] and the dimeric [Zn<sub>2</sub>(CO<sub>2</sub>)<sub>4</sub>]. The two display the same ligand disposition and distorted octahedral geometry, which seems to be stabilized by strong intermolecular N-H···O (dimer) and intramolecular O-H···O (monomer) interactions. In the monomeric unit, the  $\mu_1$ - $\eta^2$  coordination mode combined with the intramolecular interaction between the HPip and Pip ligands force geometric constraints that accommodate the metal octahedral geometry (OC-6, *S* = 2.769 (Zn1A)). By the same token, this deviation (OC-6, *S* = 3.801 (Zn1B)) [20] is emphasized in the dimer by both  $\mu_2$ - $\eta^1$ : $\eta^1$  and  $\mu_1$ - $\eta^1$ - coordination modes of the Pip ligands. Cd(II) dimers

(3 and 4) present the common  $[\text{Cd}_2(\text{CO}_2)_2]-2+2$  core usually found bearing coordination number 7. Both dimers have similar  $S$  values (2.215 and 2.448 (3) or 1.995 and 2.015 (4)) [20] since the equatorial plane arranged by  $\mu_2-\eta^2:\eta^1$  Pip linkers is almost equal. The strong double head-to-head amide  $\cdots$  amide interaction between the dimers in 4 fix the Isn ligands and amend any distortion that cannot be minimized in complex 3, in which the interactions of 4-AcPy are weaker. Overall, the smaller radius of hexacoordinated Zn(II) (0.880 Å), with respect to heptacoordinated Cd(II) (1.17 Å) [28], accentuate the geometric constraints that occur during the formation of the dimeric arrays, which is reflected in the higher  $S$  value (3.801) of complex 2 compared to complexes 3 and 4 ( $S$  range between 1.995 and 2.448) [20].

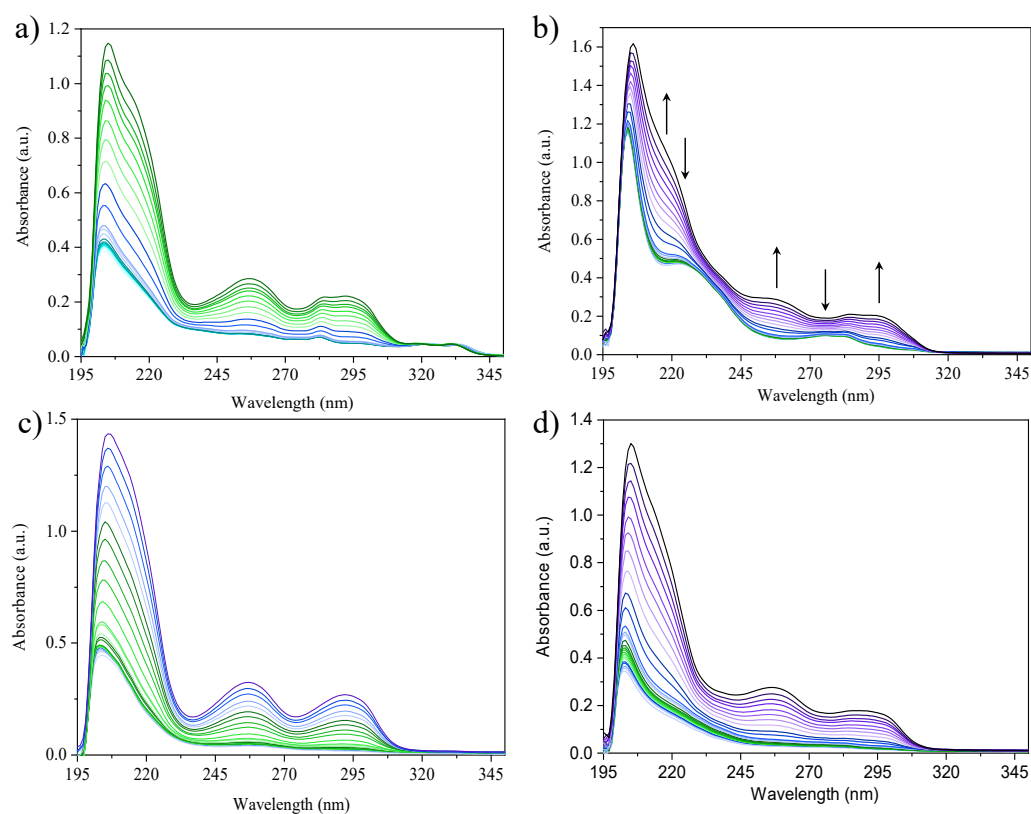


**Figure 7.** Intermolecular (a) N-H  $\cdots$  O interactions of 4. In detail, (b) amide  $\cdots$  amide interactions between A dimers or (c) between B dimers; (d) MeOH  $\cdots$  amide, (e) amide(A)  $\cdots$  Pip(B) or (f) MeOH  $\cdots$  Pip association; and (g) amide(B)  $\cdots$  Pip(A) interactions. (h)  $\pi \cdots \pi$  interactions between Pip rings.

### 2.5. Photophysical Properties

The absorption and emission properties of the complexes and ligands, as well as the references (L-tyrosine and phenanthrene), were recorded in a MeOH solution. The absorption was measured in the UV region of the spectra from 200 to 345 nm, while the emission was recorded between 270 and 450 nm at 298 K.

**UV-Vis spectroscopy.** To ensure the non-aggregation of the samples at the selected concentrations for the fluorescence experiments, we performed additive UV-Vis measurements within a concentration range from  $1 \times 10^{-9}$  to  $1 \times 10^{-5}$  M (Figure 8).



**Figure 8.** Additive UV-Vis spectra of complexes (a) **1**; (b) **2**; (c) **3** and (d) **4** performed in a concentration range from  $1 \times 10^{-9}$  M to  $1 \times 10^{-4}$  M in MeOH as the solvent at 298 K. The arrows indicate the spectral changes upon increasing concentration.

The samples do not seem to present aggregation in this range of concentration. The absorption spectrum of **2** displays a significant change from  $4.49 \times 10^{-8}$  M on, which has been ascribed to a change on the absorber rather than a change via aggregates formation. This is discussed below and supported by TD-DFT calculations. The absorption and emission maxima of complexes **1–4** ( $\lambda_{\text{max-Abs}}$  and  $\lambda_{\text{max-Em}}$ , respectively), have been identified, and their molar absorptivity ( $\epsilon$ ) and relative quantum yield ( $\phi_s$ ) calculated (Table 8).

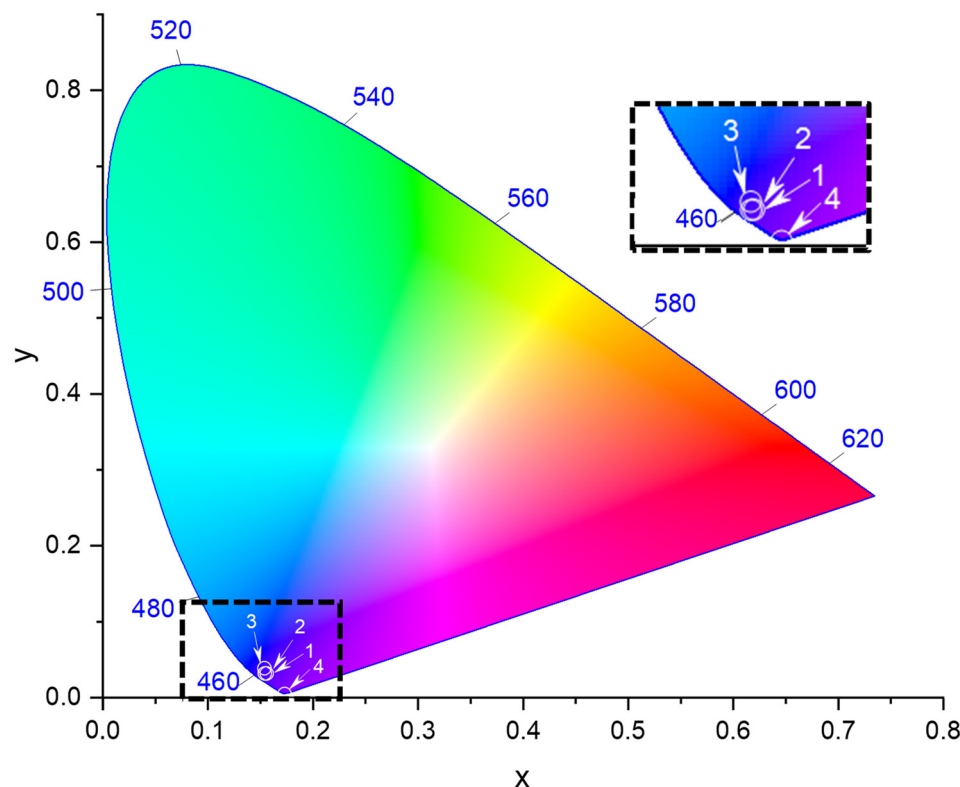
**Table 8.** UV-Vis and fluorescence data of complexes **1–4**.

Sample	$\lambda_{\text{ex}}$	$\lambda_{\text{max-em}}$	$\Delta\lambda$	$\phi_s$
<b>1</b>	263	354	9774	0.74
	225	344	15,375	0.01
<b>2</b>	225 *	344 *	15,375	0.03
	251 *	318 *	8,394	0.02
<b>3</b>	315	355	3,577	0.47
<b>4</b>	226	347	15,429	0.09

All of the wavelengths are given in nm.  $\Delta\lambda$  are given in  $\text{cm}^{-1}$ .  $\lambda_{\text{ex}}$  = excitation wavelength;  $\lambda_{\text{max-em}}$  = maximum of emission;  $\phi_s$  = quantum yield. \* Bands arising from a change in the absorber (the dimer in **2**).

**Photoluminescence.** Heretofore, fluorescence measurements were performed using concentrations of  $1.70 \times 10^{-9}$  M (**1**);  $1.08 \times 10^{-8}$  M and  $1.07 \times 10^{-7}$  M (**2**);  $1.04 \times 10^{-7}$  M (**3**); and  $1.01 \times 10^{-7}$  M (**4**), extracted from the UV-Vis results after ensuring their non-aggregation to minimize ACQ [7]; these samples were excited at the wavelength of their emission maxima. All of the relevant details have been summarized in Table 8. The spectra of **1** and **2** have single emission bands at 354 nm (**1**) and 344 or 318 nm (**2**) at being irradiated at 263 nm and 225 or 251 nm, respectively. The emission spectra of **3** and **4** present unfolded emission

bands centered at 355 and 347 nm, at being irradiated at 315 and 226 nm, respectively. As displayed in the CIE 1931 chromaticity diagram [29], the resultant emission color of 1–4 (318–355 nm) is violet (Figure 9).

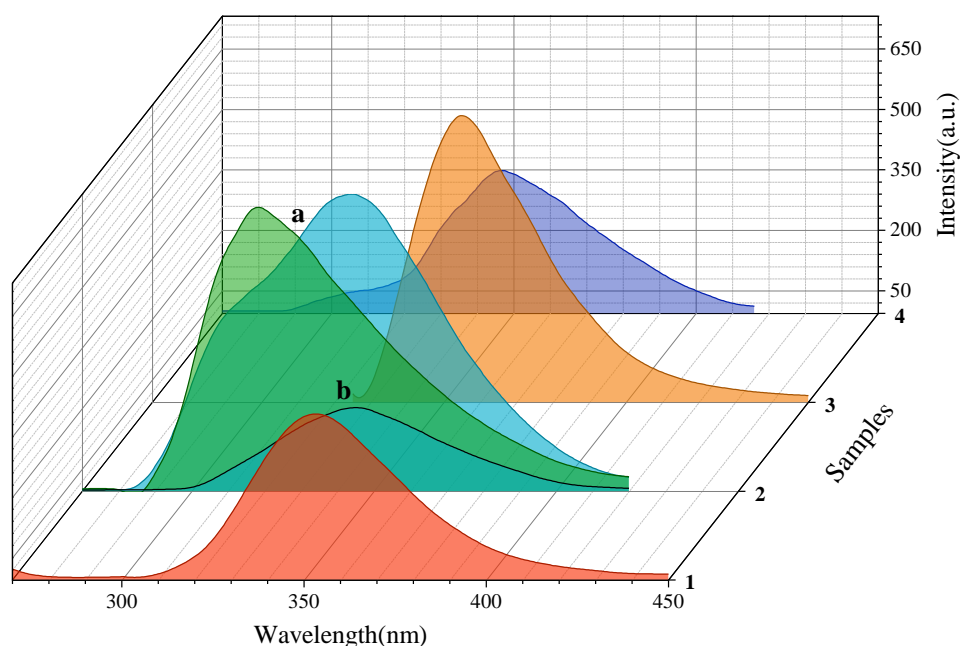


**Figure 9.** Representation of the emission color of complexes 1–4 within the CIE 1931 chromaticity diagram.

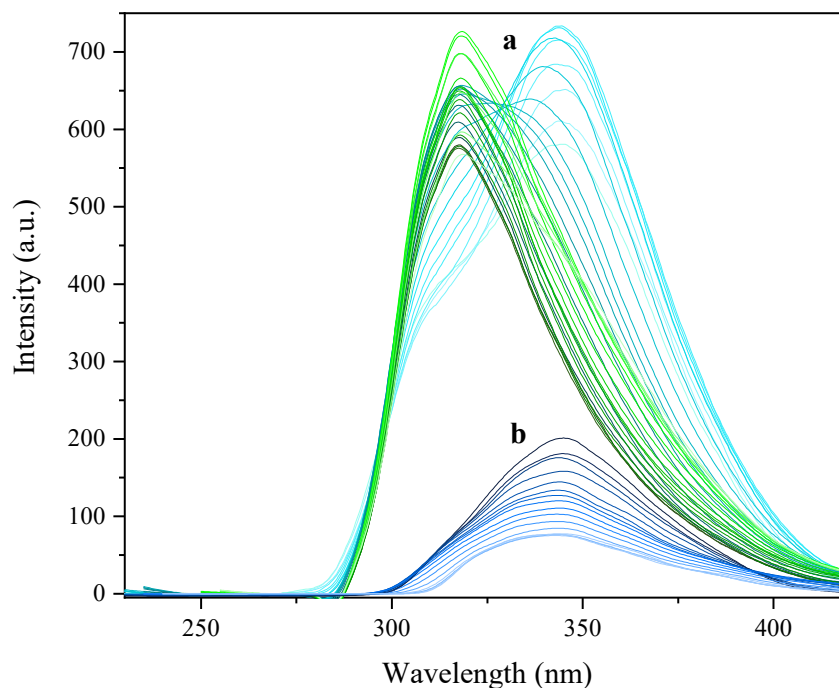
Emission intensities increase in the order  $1 < 2 < 3 < 4$  considering the different concentrations used in the fluorescence experiments (Figure 10). The relative quantum yields ( $\phi_s$ ) of the samples were calculated by way of comparison with two reference standards (L-tyrosine and phenanthrene) [30] using Equation (1):

$$\phi_s = \phi_{\text{ref}} \left( \frac{\text{OD}_{\text{ref}}}{\text{OD}_s} \right) \left( \frac{I_s}{I_{\text{ref}}} \right) \left( \frac{n_s}{n_{\text{ref}}} \right)^2 \quad (1)$$

where  $\phi$  is the quantum yield, OD is the optical density (or absorbance) at the excited wavelength,  $I$  is the area under the curve of the emission spectra, and  $n$  is the refractive index of the solvent. In this study, L-tyrosine ( $\phi_{\text{ref}} = 0.14$ ) [31] and phenanthrene ( $\phi_{\text{ref}} = 0.125$ ) [32] has been used as the standard and their  $\text{OD}_{\text{ref}}$  and  $I_{\text{ref}}$  values have been obtained from L-tyr solutions of  $1.01 \times 10^{-7}$  M in Milli-Q water ( $n_{\text{ref}} = 1.3325$ ) [33] and phenanthrene solutions of  $1.00 \times 10^{-7}$  in ethanol ( $n_{\text{ref}} = 1.3608$ ) [32] as solvents at 298K. The values of  $\phi_s$  are 0.74 (1), 0.47 (3) and 0.09 (4). Complex 2 present different quantum yields depending on the concentration and emission maxima. Under  $\lambda_{\text{ex}} = 225$  nm at  $1.08 \times 10^{-8}$  M the  $\phi_s$  value is 0.001 while at  $1 \times 10^{-7}$  M display  $\phi_s = 0.02$  under  $\lambda_{\text{ex}} = 225$  nm and 0.03 under  $\lambda_{\text{ex}} = 251$  nm (Figure 11). In addition, under exposure to black light, only complex 1 exhibits an intense yellow emission. Single crystals of 1 were irradiated under UV excitation of a pulse laser beam at  $\lambda = 326$  nm and presented an emission maxima at 559 nm with a moderate Stokes shift of  $12,786 \text{ cm}^{-1}$ . The solid state photoluminescence spectrum of 1 is shown in S.I: Figure S13.



**Figure 10.** Emission spectra in MeOH at 298 K of the samples at concentrations of  $1.70 \times 10^{-9}$  M (1); (a)  $1.07 \times 10^{-7}$  M and (b)  $1.08 \times 10^{-8}$  M (2);  $1.04 \times 10^{-7}$  M (3); and  $1.01 \times 10^{-7}$  M (4). Color codes: red (1); light green and turquoise (2); orange (3); blue (4).



**Figure 11.** Emission spectra of complex 2. (a) Sample of concentration  $1.07 \times 10^{-7}$  M irradiated from 210 to 270 nm. (b) Sample of concentration  $1.08 \times 10^{-8}$  M irradiated from 210 to 270 nm.

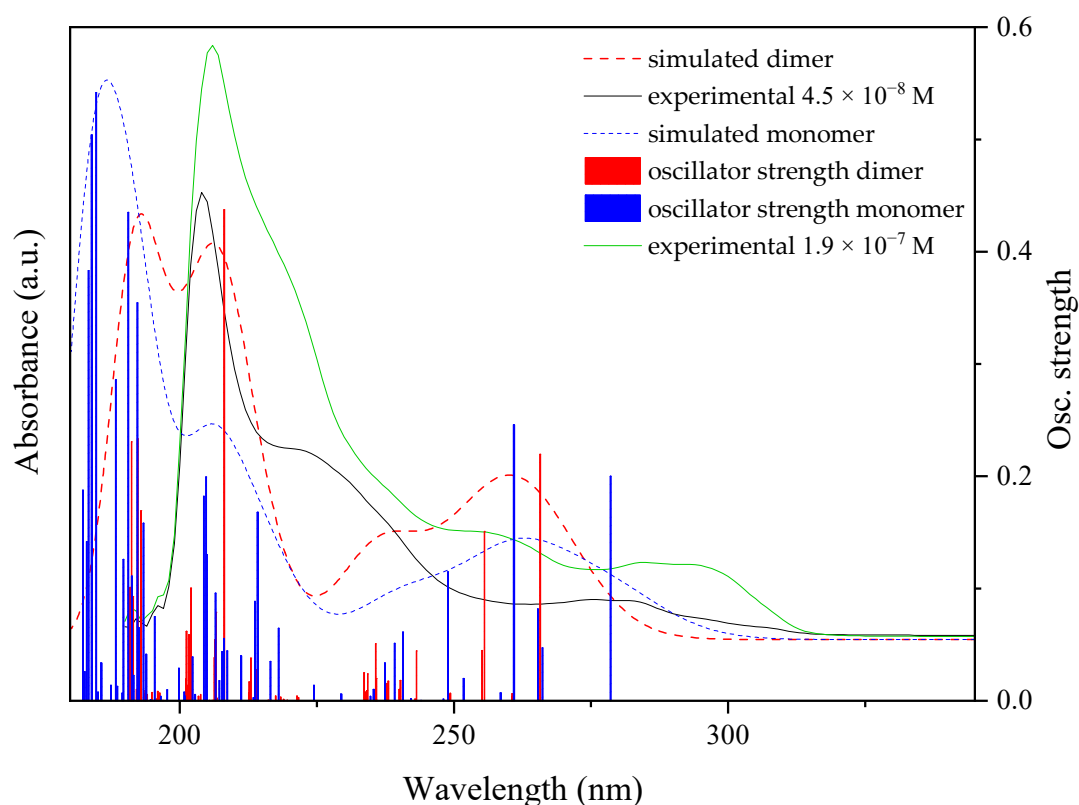
## 2.6. Electronic Calculations

**DFT calculations.** The geometric optimization of complexes 1–4 has been performed in a MeOH solution using the Polarizable Continuum Model (PCM). The resulting complexes containing 4-AcPy exhibited an improvement of the geometric evaluator  $S$  from 0.240 to 0.181 (1) and from 2.215 and 2.448 to 2.048 and 1.889 (3). In contrast, complexes 2 and 4 presented significant differences between the X-ray diffraction data and the geometry in MeOH solution. The monomeric array in compound 2 is already stabilized by strong



O-H···O intramolecular interactions between HPip and Pip ligands and suffers minor geometric changes with  $S$  values varying from 2.769 to 2.311. However, the dimer in **2**, which is stacked in chains by strong amide···amide intermolecular interactions, presents significant differences between when they are in a solid state compared to being in solution. The  $S$  values change from 3.801, which corresponds to an octahedral geometry, to 8.571, acquiring a trigonal prismatic geometry (S.I: Table S2). This geometrical change can be attributed to the strong intramolecular amide···amide interactions between the stacked Isn ligands that force the geometry of the Zn(II) metal nodes. A similar difference between the solid state and the MeOH solution is also present in complex **4**; however, in this case, the bigger size of Cd(II) allows for a structural reorganization without such a marked difference in the  $S$  values, which shift from 2.015 and 1.995 to 3.367 and 2.288. Since MeOH has a solvent polarity parameter [34] ET(30) of  $55.4 \text{ kcal}\cdot\text{mol}^{-1}$ , this could be insufficient to predominantly establish the intermolecular interactions with Isn and prevent the intramolecular amide···amide association. Therefore, the dimers in complexes **2** and **4** are amenable to experience significant geometric changes and promote relaxation through non-radiative decays.

**TD-DFT calculations.** All of the calculated UV-Vis spectra of complexes **1–4** agree reasonably well with the experimental profiles (Figure 12, S.I.: Figures S14–S16). The shift in the theoretical absorption spectra with respect to the experimental profiles is within the range of typical TD-DFT calculations ( $\sim 0.3 \text{ eV}$ ) and are caused by computing the absorptions as vertical transitions [35]. Only transitions with a higher oscillator strength ( $f$ ) value have been selected for the molecular orbitals representation and natural transition orbitals (NTOs) analysis. The HOMO and LUMO outline as well as the energy gaps can be found in the S.I: Figure S17. Subsequently, the main contributors of the electronic transitions have been analyzed for each absorption band to identify the regions involved in it. The molecular orbitals of each set of transitions have been represented, as well as the corresponding NTOs.



**Figure 12.** Experimental (black and green lines), calculated (dashed black and green lines) UV-Vis spectra and oscillator strengths of  $[\text{Zn}(\mu\text{-Pip})(\text{Pip})(\text{Isn})_2]_2 \cdot 2[\text{Zn}(\text{Pip})_2(\text{HPip})(\text{Isn})] \cdot 2\text{MeOH}$  (**2**).

The TD-DFT results of the monomeric and dimeric forms present in complex **2** resulted in the ascription of the spectrum obtained at concentrations below  $4.49 \times 10^{-8}$  M to the monomeric form, while the shape of the spectra of the dimer is closer to the one resulting from higher concentrations. Therefore, from the electronic calculations, it could be stated that the monomer–dimer ratio is displaced towards the monomeric form at lower concentrations while a mixture of both or even primacy of the dimer is observed.

The HOMO and LUMO orbitals of **1–4** have  $\pi$  symmetry, being the HOMO along the Pip ligand while the LUMO is localized over the 4-Acpy and Isn linkers. This MO seclusion was previously observed [15] suggesting that by keeping the Pip linker constant, the incorporation of 4-Acpy ligand, bearing a more electron withdrawing functional group, would lower the energy of the LUMO and, thus, reduce the energy gap. The HOMO orbitals of these complexes are quite similar in energy, ranging from  $-8.312$  eV to  $-8.018$  eV since they are all located over the Pip ligands. In opposite, the LUMO orbitals present a significant difference in energy, which can be sorted into three groups. The monomer in **2** has the higher energy LUMO of  $-0.038$  eV, while the dimers of Isn have values of  $-0.407$  eV (**2**) and  $-0.535$  eV (**4**), respectively. Finally, **1** and **3** dimers have the lowest energy LUMO with values of  $-0.905$  eV (**1**) and  $-0.976$  eV (**3**). The HOMO–LUMO gaps are between  $8.274$  eV and  $7.244$  eV, with complexes **1** and **3** presenting the smallest values.

The combined MOs (S.I.: Figures S18–S22) and NTOs (S.I.: Figures S23–S27) analysis of **1–4** reveals that less energetic electronic transitions of the spectra are based on local excitations (LE) and are mainly centered on the Pip ligands. Then, between 240 and 270 nm complexes **1** and **3** present LE over 4-Acpy, while **2** and **4** tend to promote charge transfer transitions between Pip and Isn and also have small contribution of MLCT to Isn. At higher energy, all but complex **1** present MLCT transitions from Pip to 4-Acpy or Isn. Since complex **1** does not present MLCT nor LMCT transitions, there appears to be a structural effect of the paddle-wheel that, in this case, minimizes electronic transitions on the Zn(II) metal center and hinders the charge transfers between Pip and 4-Acpy. The spatial arrangement of the ligands forced by the paddle-wheel leads to a greater separation between them, thus, avoiding intramolecular charge transfer transitions. The electronic transition states (TS) at the selected  $\lambda_{\text{ex}}$  of complexes containing Isn (**2** and **4**) present a combination of LE over Isn with a strong contribution of ILCT transitions and a small contribution of MLCT to Isn in **4** (TS20 of the monomer, TS41 of the dimer in **2** or TS23 and 24 of **4**); whereas, 4-Acpy complexes mainly present LE over 4-Acpy (TS8 and 19 of **1**) or ILCT between them (TS5 and 6 of **3**). Therefore, the combination of both intramolecular interactions and dimeric structure could directly affect the photophysical properties by promoting charge transfers instead of LE; emphasized in Isn complexes. The Supporting Information displays the complete data about geometry optimization (S.I.: Tables S3–S7 and Figures S28–S32).

### 3. Conclusions

A series of Zn(II) and Cd(II) complexes with HPip, 4-Acpy and Isn have been synthesized and fully characterized. Their crystal structure consists of one Zn(II) dimeric paddle-wheel (**1**); a mixture of one dimer and two monomers in the unit cell (**2**); and two dimers of Cd(II) (**3** and **4**). Their different nuclearity is based on different combinations of coordination modes of the Pip ligand: monodentate ( $\mu_1\text{-}\eta^1$ , **2**), bidentate chelate ( $\mu_1\text{-}\eta^2$ , **2–4**), bridged ( $\mu_2\text{-}\eta^1:\eta^1$ , **1** and **2**) or both ( $\mu_2\text{-}\eta^2:\eta^1$ , **3** and **4**) strongly influenced by intra- and intermolecular interactions. In them, Zn(II) metal node displays coordination numbers of 5 (**1**) and 6 (**2**) while Cd(II) exhibits coordination number 7 (**3** and **4**). DFT geometric optimizations revealed dPy dependent geometrical changes, emphasized by the formation of intramolecular amide...amide interactions. TD-DFT results show HOMO–LUMO gap dependence on the dPy, being the shortest corresponding to the 4-Acpy complexes. NTOs analysis revealed LE character of absorptions in **1** and **3**; whereas, in **2** and **4**, ILCTs have a significant contribution. This has also been reflected in the  $\phi_S$  results in complex **1** being up to ten times higher than in **2**. These results reflect how (i) hampering charge transfer

transition by avoiding intramolecular  $\pi \cdots \pi$  interactions; (ii) minimizing PET processes through coordination to Zn(II) and Cd(II); and (iii) changes in the *p*-substituents of dPy, can modulate the HOMO–LUMO gap and maximize the resulting  $\sigma_s$  values.

## 4. Experimental

### 4.1. Materials and General Details

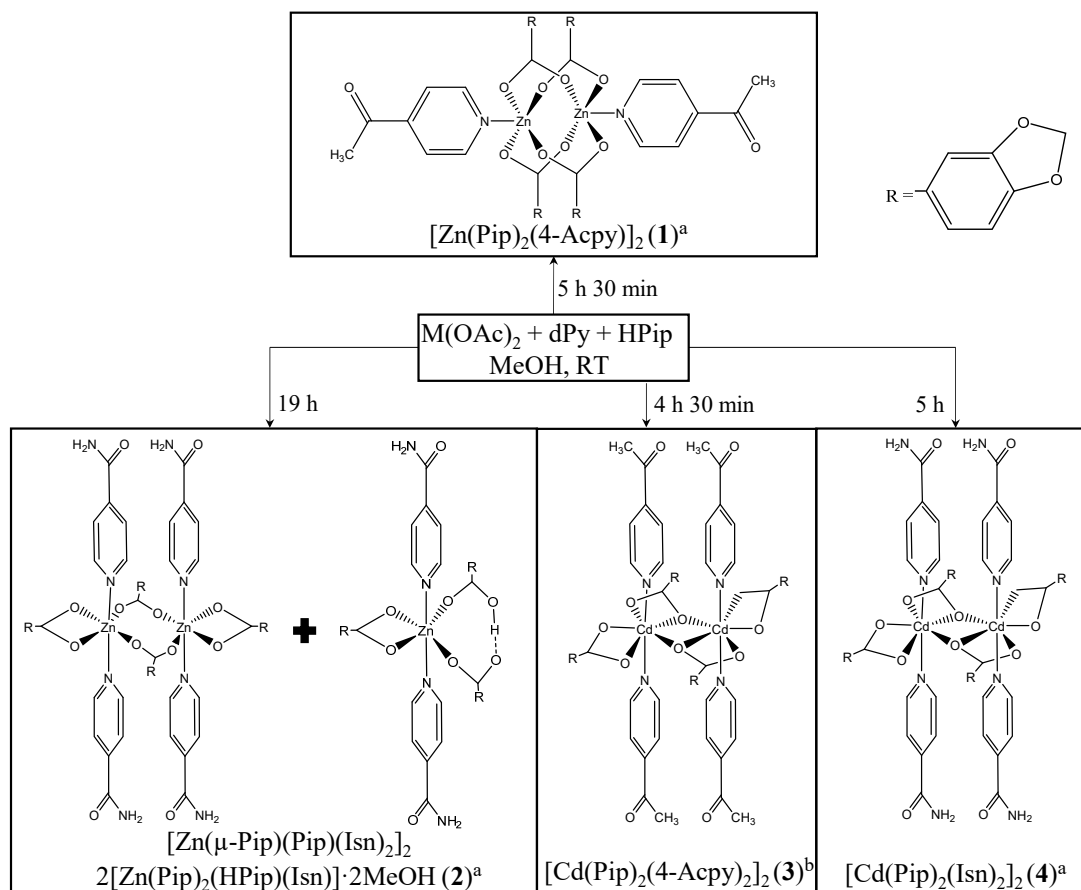
Zinc(II) acetate dihydrate ( $\text{Zn}(\text{OAc})_2 \cdot 2\text{H}_2\text{O}$ ), cadmium(II) acetate dihydrate ( $\text{Cd}(\text{OAc})_2 \cdot 2\text{H}_2\text{O}$ ), 1,3-benzodioxole-5-carboxylic acid (Piperonylic, HPip), 4-acetylpyridine (4-AcPy) and isonicotinamide (Isn) ligands; and methanol (MeOH), ethanol (EtOH) and diethyl ether ( $\text{Et}_2\text{O}$ ) as solvents, were purchased from Sigma-Aldrich (Sigma Aldrich, St. Louis, MO, USA). L-tyrosine and phenanthrene were also purchased from Sigma-Aldrich, and used as the reference standard for fluorescence measurements. Deuterated dimethyl sulfoxide ( $\text{dmsol-}d_6$ ) was purchased from Eurisotop (Euriso-Top GmbH, Saarbrücken, Germany). All of them were used without further purification. All reactions and manipulation were carried out in air at RT (Scheme 1). The elemental analyses (C, H, N) were carried out on a Euro Vector 3100 instrument (Eurovector, Pavia, Italy). The FTIR-ATR spectra were recorded on a Perkin Elmer spectrometer (Perkin Elmer Inc, Shelton, CT, USA), equipped with a universal attenuated total reflectance (ATR) accessory with a diamond window in the range of  $4000\text{--}500\text{ cm}^{-1}$ .  $^1\text{H}$ ,  $^{13}\text{C}\{^1\text{H}\}$  and DEPT-135 NMR spectra were recorded on an NMR-FT Bruker360 MHz spectrometer (Bruker BioSpin MRI GmbH, Ettlingen, Germany) spectrometer in  $\text{DMSO-}d_6$  solution at RT. All chemical shifts ( $\delta$ ) are given in ppm. The electronic spectra in MeOH solution were run on an Agilent HP 8453 UV-Vis spectrophotometer (Agilent, Santa Clara, CA, USA) with a quartz cell having a path length of 1 cm in the range of 190–345 nm. The molar absorptivity values were calculated and displayed as  $\log(\epsilon)$ . The fluorescence measurements were carried out at  $25\text{ }^\circ\text{C}$  with a PerkinElmer LS 55 50 Hz fluorescence spectrometer (Perkin Elmer Inc., Shelton, CT, USA) using a 1 cm quartz cell, in MeOH solution. The samples were excited at their absorption maxima and the emission was recorded between 195 and 450 nm. The solid state photoluminescence measurement was recorded between 500 and 630 nm using a Varian Cary Eclipse Fluorescence spectrophotometer (Agilent, Santa Clara, CA, USA) and is given in nm. The UV-Vis and fluorescence spectra in solution, and the spectrum of **1** in solid state, were measured under air exposure. Both CIE 1931 chromaticity diagrams as well as the corrected dilution effects on the UV-Vis and fluorescence data were performed by means of Origin Pro 2019b software (OriginLab Corporation, Northampton, MA, USA).

### 4.2. Synthesis of Complexes 1–4

$[\text{Zn}(\mu\text{-Pip})_2(4\text{-AcPy})_2]$  (**1**). To a MeOH solution (15 mL) of  $\text{Zn}(\text{OAc})_2 \cdot 2\text{H}_2\text{O}$  (250 mg, 1.14 mmol), the 4-AcPy (0.500 mL, 4.13 mmol) was added drop wise, and the mixture was stirred for 5 min. Then, a MeOH solution (35 mL) of HPip (379 mg, 2.28 mmol) was added drop wise. The resulting solution was stirred for 5 h 30 min until a yellowish powder precipitated. The suspension was cooled down for 15 min and filtered off. The solid was washed twice with cold MeOH (5 mL) and dried under vacuum. Suitable colorless crystals were obtained via slow evaporation of the mother liquors for 6 days.

Yield: 365 mg (62%) (respect to  $\text{Zn}(\text{OAc})_2 \cdot 2\text{H}_2\text{O}$ ). Elem. Anal. Calc. for  $\text{C}_{46}\text{H}_{34}\text{Zn}_2\text{N}_2\text{O}_{18}$  ( $1033.54\text{ g}\cdot\text{mol}^{-1}$ ): C 53.46; H 3.32; N 2.71. Found C 53.28; H 3.05; N 2.64. FTIR-ATR (wavenumber,  $\text{cm}^{-1}$ ): 3107(w)–3001 [ $\nu_{\text{ar}}(\text{C-H})$ ], 2912(w) [ $\nu_{\text{al}}(\text{C-H})$ ], 1703(m) [ $\nu(\text{C=O})$ ], 1638(m) [ $\nu(\text{C=C})$ ,  $\nu(\text{C=N})$ ], 1597(m) [ $\nu(\text{C=C})$ ,  $\nu(\text{C=N})$ ], 1556(m) [ $\nu_{\text{as}}(\text{COO})$ ], 1502(m), 1437(s), 1385(s) [ $\nu_{\text{s}}(\text{COO})$ ], 1355(m) [ $\delta(\text{C=C})$ ,  $\delta(\text{C=N})$ ], 1255(s), 1237(m), 1213(m), 1167(m) [ $\nu(\text{C-O-C})$ ], 1107(m), 1078(w), 1025(s) [ $\delta_{\text{ip}}(\text{C-H})$ ], 930(w), 919(m), 883(m), 807(m), 771(s) [ $\delta_{\text{oop}}(\text{C-H})$ ], 730(m), 677(m), 588(m).  $^1\text{H}$  NMR (400 MHz;  $\text{DMSO-}d_6$ ; 298 K):  $\delta$  = 8.71 [2H, d,  $^3\text{J}$  = 4.1 Hz, *o-H*] $_{4\text{-AcPy}}$ , 7.72 [2H, d,  $^3\text{J}$  = 4.1 Hz, *m-H*] $_{4\text{-AcPy}}$ , 7.43 [2H, dd,  $^3\text{J}$  = 8.0 Hz,  $^4\text{J}$  = 1.6 Hz,  $\text{O}_2\text{C-CH-CH}$ ], 7.27 [2H, d,  $^4\text{J}$  = 1.6 Hz,  $\text{O}_2\text{C-CH-CO}$ ], 6.80 [2H, d,  $^3\text{J}$  = 7.9 Hz,  $\text{O}_2\text{C-CH-CH}$ ], 5.95 [4H, s,  $\text{O-CH}_2\text{-O}$ ], 2.39 [3H, s,  $-\text{CH}_3$ ].  $^{13}\text{C}\{^1\text{H}\}$  NMR (400 MHz;  $\text{DMSO-}d_6$ ; 298 K):  $\delta$  =

198.1 [OC-CH<sub>3</sub>], 171.4 [O<sub>2</sub>C-C], 150.8 [N-CH], 149.8 [O<sub>2</sub>C-C-(CH)<sub>2</sub>-C], 147.1 [O<sub>2</sub>C-C-CH-C], 143.1 [N-(CH)<sub>2</sub>-C], 128.9 [O<sub>2</sub>C-C], 124.7 [N-CH-CH], 121.8 [O<sub>2</sub>C-C-CH-CH], 109.5 [O<sub>2</sub>C-C-CH-C], 107.7 [O<sub>2</sub>C-C-CH-CH], 101.6 [O-CH<sub>2</sub>-O], 27.0 [-CH<sub>3</sub>]. UV-Vis in MeOH; wavelength (ε, M<sup>-1</sup>·cm<sup>-1</sup>): 203(4.93); 258(4.40); 282(4.20); 292(4.30).



**Scheme 1.** Reactions for synthesis of metal complexes **1**, **2** and **4** were performed by adding a methanolic solution of HPip over a mixture of  $M(OAc)_2$  and 4-AcPy or Isn<sup>a</sup>. The synthesis of complex **3** was achieved by adding a mixture of 4-AcPy and HPip to a methanolic solution of  $Cd(OAc)_2 \cdot 2H_2O$ <sup>b</sup>.

$[Zn(\mu-Pip)(Pip)(Isn)_2]_2 \cdot 2[Zn(Pip)_2(HPip)(Isn)] \cdot 2MeOH (2)$ . To a MeOH solution (2 mL) of  $Zn(OAc)_2 \cdot 2H_2O$  (74.9 mg, 0.341 mmol), a MeOH solution (6 mL) of Isn ligand (83.1 mg, 0.680 mmol) was added drop wise and the mixture was stirred for 5 min. Subsequently, a MeOH solution (16 mL) of HPip (114 mg, 0.682 mmol) was added drop wise. The solution was stirred for 4 h 30 min. Then, the reaction was concentrated until it reached half of the original volume, before being cooled down in an ice bath until a white powder precipitated. The resulting solid was filtered off, washed twice with cold Et<sub>2</sub>O (5 mL) and dried under vacuum. Suitable colorless crystals were obtained via slow evaporation of the mother liquors for 3 days.

Yield: 208 mg (83%) (respect to  $Zn(OAc)_2 \cdot 2H_2O$ ). Elem. Anal. Calc. for  $C_{130}H_{108}Zn_4N_{16}O_{50}$  (2955.80 g·mol<sup>-1</sup>): C 52.82; H 3.68; N 7.58. Found C 52.73; H 3.52; N 7.50. FTIR-ATR (wavenumber, cm<sup>-1</sup>): 3445(w) [ $\nu(O-H)_{MeOH}$ ], 3318(w) [ $\nu_{as}(N-H)$ ], 3172(w) [ $\nu_s(N-H)$ ], 3111(w)–3071 (w) [ $\nu_{ar}(C-H)$ ], 2904(w)–2789(w) [ $\nu_{al}(C-H)$ ], 1708(m) [ $\nu(C=O)_{Isn}$ ], 1678(m) [ $\nu(COOH)$ ], 1629(m) [ $\nu(C=C)$ ,  $\nu(C=N)$ ], 1607(w) [ $\nu(C=C)$ ,  $\nu(C=N)$ ], 1553(m), 1534(m) [ $\nu_{as}(COO)$ ], 1502(m) [ $\nu_{as}(COO)$ ], 1435(s), 1417(s), 1385(s) [ $\nu_s(COO)$ ], 1353(s), 1300(m) [ $\delta(C=C)$ ,  $\delta(C=N)$ ], 1256(s), 1240(s), 1170(m) [ $\nu(C-O-C)$ ], 1147(w), 1113(m), 1075(w), 1063(w), 1035(s) [ $\delta_{ip}(C-H)$ ], 944(m), 928(m), 876(w), 861(m), 831(m), 804(m), 773(s) [ $\delta_{oop}(C-H)$ ],

720(m), 681 (m), 641(m), 580(m), 531(m).  $^1\text{H}$  NMR (360 MHz; DMSO- $d_6$ ; 298 K):  $\delta$  = 8.66 [3.2H, d,  $^3J$  = 3.9 Hz, *o*-H]<sub>Isn</sub>, 8.18 [1.6H, s, OC-NH], 7.72 [3.1H, d,  $^3J$  = 4.7 Hz, N-CH-CH], 7.66 [1.6H, s, OC-NH], 7.47 [2H, d,  $^3J$  = 8.1 Hz, O<sub>2</sub>C-C-CH-CH], 7.30 [2H, s, O<sub>2</sub>C-C-CH-CO], 6.86 [2H, d,  $^3J$  = 7.5 Hz, O<sub>2</sub>C-C-CH-CH], 6.00 [4H, s, O-CH<sub>2</sub>-O], 4.03 [0.4H, br., -OH], 3.10 [1.2H, s, -CH<sub>3</sub>].  $^{13}\text{C}\{^1\text{H}\}$  NMR (360 MHz; DMSO- $d_6$ ; 298 K):  $\delta$  = 171.2 [O<sub>2</sub>C-C], 166.3 [OC-NH<sub>2</sub>], 150.2 [O<sub>2</sub>C-C-(CH)<sub>2</sub>-C], 149.6 [N-CH], 146.9 [O<sub>2</sub>C-C-CH-C], 141.6 [N-(CH)<sub>2</sub>-C], 128.9 [O<sub>2</sub>C-C], 124.6 [N-CH-CH], 121.6 [O<sub>2</sub>C-C-CH-CH], 109.4 [O<sub>2</sub>C-C-CH-C], 107.5 [O<sub>2</sub>C-C-CH-CH], 101.4 [O-CH<sub>2</sub>-O]. UV-Vis in MeOH; wavelength ( $\epsilon$ , M<sup>-1</sup>·cm<sup>-1</sup>): 204(4.59); 225(3.24); 251(4.32); 276(4.05); 285(4.13); 292(4.26).

[Cd( $\mu$ -Pip)(Pip)(4-Acpy)<sub>2</sub>]<sub>2</sub> (3). To a MeOH solution (20 mL) of Cd(OAc)<sub>2</sub>·2H<sub>2</sub>O (100 mg, 0.375 mmol), a MeOH solution (30 mL) of HPip (125 mg, 0.750 mmol) and 4-Acpy (179 mg, 1.47 mmol) was added drop wise. The resulting solution was stirred for 19 h, concentrated to half of the original volume under vacuum, and kept in the fridge for 2 days until a white precipitate was formed. The powder was filtered, washed twice with cold MeOH (10 mL) and dried under vacuum. Suitable colorless crystals were obtained via slow evaporation of the mother liquors for 8 days.

Yield: 141 mg (55%) (respect to Cd(OAc)<sub>2</sub>·2H<sub>2</sub>O). Elem. Anal. Calc. for C<sub>60</sub>H<sub>48</sub>Cd<sub>2</sub>N<sub>4</sub>O<sub>20</sub> (1369.86 g·mol<sup>-1</sup>): C 52.61; H 3.53; N 4.09. Found C 52.34; H 3.47; N 3.95. FTIR-ATR (wavenumber, cm<sup>-1</sup>): 3066(w) [ $\nu_{\text{ar}}$ (C-H)], 3034(w) [ $\nu_{\text{ar}}$ (C-H)], 2905(m)–2779(w) [ $\nu_{\text{al}}$ (C-H)], 1693(m) [ $\nu$ (C=O)], 1632(w) [ $\nu$ (C=C),  $\nu$ (C=N)], 1603(w) [ $\nu$ (C=C),  $\nu$ (C=N)], 1539(m) [ $\nu_{\text{as}}$ (COO)], 1506(m) [ $\nu_{\text{as}}$ (COO)], 1436(m), 1415(m), 1384(s) [ $\nu_{\text{s}}$ (COO)], 1360(s) [ $\delta$ (C=C),  $\delta$ (C=N)], 1254(m), 1170(w), 1113(m) [ $\nu$ (C-O-C)], 1035(m) [ $\delta_{\text{ip}}$ (C-H)], 1014(m), 923(m), 822(m) [ $\delta_{\text{oop}}$ (C-H)], 806(m) [ $\delta_{\text{oop}}$ (C-H)], 775(s) [ $\delta_{\text{oop}}$ (C-H)], 682(m), 588(m), 530(s).  $^1\text{H}$  NMR (400 MHz; DMSO- $d_6$ ; 298 K):  $\delta$  = 8.82 [4H, br, *o*-H]<sub>4-Acpy</sub>, 7.82 [4H, dd,  $^3J$  = 6.0,  $^4J$  = 2.8 Hz, *m*-H]<sub>4-Acpy</sub>, 7.55 [2H, dd,  $^3J$  = 8.1 Hz,  $^4J$  = 1.6 Hz, O<sub>2</sub>C-CH-CH], 7.40 [2H, d,  $^4J$  = 1.4 Hz, O<sub>2</sub>C-CH-CO], 6.92 [2H, d,  $^3J$  = 8.1 Hz, O<sub>2</sub>C-CH-CH], 6.07 [4H, s, O-CH<sub>2</sub>-O], 2.63 [6H, s, -CH<sub>3</sub>].  $^{13}\text{C}\{^1\text{H}\}$  NMR (360 MHz; DMSO- $d_6$ ; 298 K):  $\delta$  = 198.5 [OC-CH<sub>3</sub>], 172.3 [O<sub>2</sub>C-C], 151.2 [N-CH], 149.9 [O<sub>2</sub>C-C-(CH)<sub>2</sub>-C], 147.3 [O<sub>2</sub>C-C-CH-C], 143.1 [N-(CH)<sub>2</sub>-C], 129.5 [O<sub>2</sub>C-C], 125.1 [N-CH-CH], 121.9 [O<sub>2</sub>C-C-CH-CH], 110.0 [O<sub>2</sub>C-C-CH-C], 107.9 [O<sub>2</sub>C-C-CH-CH], 101.8 [O-CH<sub>2</sub>-O], 27.3 [-CH<sub>3</sub>]. UV-Vis in MeOH; wavelength ( $\epsilon$ , M<sup>-1</sup>·cm<sup>-1</sup>): 203(4.95); 257(4.51); 291(4.45).

[Cd( $\mu$ -Pip)<sub>2</sub>(Isn)<sub>2</sub>]<sub>2</sub>·MeOH (4). To a MeOH solution (2 mL) of Cd(OAc)<sub>2</sub>·2H<sub>2</sub>O (150 mg, 0.651 mmol), a solution (6 mL) of Isn ligand (159 mg, 1.30 mmol) was added drop wise and the mixture was stirred for 5 min. Subsequently, a MeOH solution (16 mL) of HPip (226 mg, 1.36 mmol) was added drop wise. After 5 min of stirring, a white powder precipitated. The reaction was kept under stirring conditions for 5 h. Then, it was cooled down in an ice bath and the resulting solid was filtered off, washed twice with cold Et<sub>2</sub>O (5 mL) and dried under vacuum. Suitable colorless crystals were obtained via recrystallization in MeOH.

Yield: 325 mg (72%) (respect to Cd(OAc)<sub>2</sub>·2H<sub>2</sub>O). Elem. Anal. Calc. for C<sub>57</sub>H<sub>48</sub>Cd<sub>2</sub>N<sub>8</sub>O<sub>21</sub> (1405.85 g·mol<sup>-1</sup>): C 48.70; H 3.44; N 7.97. Found C 48.75; H 3.41; N 7.88. FTIR-ATR (wavenumber, cm<sup>-1</sup>): 3442(w) [ $\nu$ (N-H)], 3316(w) [ $\nu$ (N-H)], 3171(m) [ $\nu$ (N-H)], 3069(w) [ $\nu_{\text{ar}}$ (C-H)], 2972(w)–2900(w) [ $\nu_{\text{al}}$ (C-H)], 1704(m) [ $\nu$ (C=O)<sub>Isn</sub>], 1626(m) [ $\nu$ (C=C),  $\nu$ (C=N)], 1611(w) [ $\nu$ (C=C),  $\nu$ (C=N)], 1553(sh.), 1539(s) [ $\nu_{\text{as}}$ (COO)], 1502(m) [ $\nu_{\text{as}}$ (COO)], 1434(s), 1415(m), 1385(s) [ $\nu_{\text{s}}$ (COO)], 1349(m) [ $\delta$ (C=C),  $\delta$ (C=N)], 1256(s), 1237(m), 1168(m) [ $\nu$ (C-O-C)], 1147(w), 1108(m), 1066(w), 1039(s) [ $\delta_{\text{ip}}$ (C-H)], 1015(m), 942(w), 921(m), 885(w), 855(m), 825(m), 806(m), 777(s) [ $\delta_{\text{oop}}$ (C-H)], 723(m), 684(m), 641(s), 581(m), 530(m).  $^1\text{H}$  NMR (360 MHz; DMSO- $d_6$ ; 298 K):  $\delta$  = 8.71 [4H, d,  $^3J$  = 4.1 Hz, *o*-H]<sub>Isn</sub>, 8.23 [2H, s, OC-NH]<sub>Isn</sub>, 7.76 [2H, dd,  $^3J$  = 4.4 Hz,  $^4J$  = 1.8 Hz, N-CH-CH], 7.72 [2H, s, OC-NH]<sub>Isn</sub>, 7.54 [2H, d,  $^3J$  = 8.0 Hz, O<sub>2</sub>C-C-CH-CH], 7.39 [2H, s, O<sub>2</sub>C-C-CH-CO], 6.91 [2H, d,  $^3J$  = 8.1 Hz, O<sub>2</sub>C-C-CH-CH], 6.06 [4H, s, O-CH<sub>2</sub>-O].  $^{13}\text{C}\{^1\text{H}\}$  NMR (360 MHz; DMSO- $d_6$ ; 298 K):  $\delta$  = 172.1



[O<sub>2</sub>C-C], 166.6 [OCNH<sub>2</sub>], 150.5 [O<sub>2</sub>C-C-(CH)<sub>2</sub>-C], 149.7 [N-CH], 147.1 [O<sub>2</sub>C-C-CH-C], 141.6 [N-(CH)<sub>2</sub>-C], 141.1 [O<sub>2</sub>C-C], 129.2 [N-CH-CH], 124.9 [O<sub>2</sub>C-C], 121.8 [O<sub>2</sub>C-C-CH-CH], 109.8 [O<sub>2</sub>C-C-CH-C], 107.7 [O<sub>2</sub>C-C-CH-CH], 101.6 [O-CH<sub>2</sub>-O]. UV-Vis in MeOH; wavelength ( $\epsilon$ , M<sup>-1</sup>·cm<sup>-1</sup>): 203(4.94); 256(4.41); 285(4.24).

#### 4.3. X-ray Crystallography

Colorless (**1**, **2** and **4**) and yellow (**3**) prism-like specimens were used for the X-ray crystallographic analysis. The X-ray intensity data were measured on a D8 Venture system (Bruker, Karlsruhe, Germany) equipped with a multilayer monochromator and a Mo microfocus ( $\lambda = 0.71073 \text{ \AA}$ ). For **1–4**, the frames were integrated with the Bruker SAINT software package (Bruker, Karlsruhe, Germany) using a narrow-frame algorithm. For **1**, the integration of the data yielded 8,790 independent reflections (average redundancy 5.448,  $R_{\text{sig}} = 2.65\%$ ) and 7,734 (87.99%) were greater than  $2\sigma(|F|^2)$ . The root-mean-square (RMS) deviation in the largest hole was  $0.092 \text{ e}^-/\text{\AA}^3$ . For **2**, the integration of the data yielded 18,771 independent reflections (average redundancy 5.125,  $R_{\text{sig}} = 8.69\%$ ) and 10,867 (57.89%) were greater than  $2\sigma(|F|^2)$ . The RMS deviation in the largest hole was  $0.127 \text{ e}^-/\text{\AA}^3$ . For **3**, the integration of the data yielded 12,559 independent reflections (average redundancy 6.025,  $R_{\text{sig}} = 4.93\%$ ) and 9725 (77.43%) were greater than  $2\sigma(|F|^2)$ . The RMS deviation in the largest hole was  $-1.028 \text{ e}^-/\text{\AA}^3$ . For **4**, the integration of the data yielded 18,818 independent reflections (average redundancy 1.000,  $R_{\text{sig}} = 4.87\%$ ) and 12,418 (65.99%) were greater than  $2\sigma(|F|^2)$ . The RMS deviation in the largest hole was  $0.123 \text{ e}^-/\text{\AA}^3$ .

The structures of **1–4** were solved and refined using the Bruker SHELXTL Software Package (version-2018/3) [36]. The final cell constants and volume are based upon the refinement of the XYZ-centroids of reflections above  $20 \sigma(I)$ . The data were corrected for absorption effects using the multi-scan method (SADABS); the crystal data and relevant details of the structure refinement for compounds **1–4** are reported in Tables 9 and 10. Complete information about the crystal structure and molecular geometry is available in .cif format as Supporting Information: CCDC 2124893 (**1**), 2124895 (**2**), 2124894 (**3**) and 2124892 (**4**). The molecular graphics were generated with Mercury 4.2.0 software [37] using the POV-Ray image package [38]. The color codes for all molecular graphics are: blue (Zn), yellow (Cd), light blue (N), red (O), grey (C), and white (H).

**Table 9.** Crystal structure refinement parameters for **1** and **2**.

	<b>1</b>	<b>2</b>
Empirical formula	C <sub>46</sub> H <sub>34</sub> ZnN <sub>2</sub> O <sub>18</sub>	C <sub>130</sub> H <sub>108</sub> Zn <sub>4</sub> N <sub>16</sub> O <sub>50</sub>
Formula weight	1033.49	2955.80
T (K)	100(2)	100(2)
Wavelength (Å)	0.71073	0.71073
System, space group	Triclinic, P-1	Triclinic, P-1
Unit cell dimensions		
a (Å)	8.2648(5)	10.8798(6)
b (Å)	11.5381(6)	16.9461(8)
c (Å)	11.7350(7)	17.3401(9)
$\alpha$ (°)	103.454(2)	84.455(2)
$\beta$ (°)	90.849(2)	80.089(2)
$\gamma$ (°)	105.143(2)	77.457(2)
V (Å <sup>3</sup> )	1047.14(11)	3068.3(3)
Z	1	1

Table 9. Cont.

	1	2
$D_{\text{calc}}$ (g cm <sup>3</sup> )	1.639	1.600
$\mu$ (mm <sup>-1</sup> )	1.231	0.878
$F(000)$	528	1520
Crystal size (mm <sup>3</sup> )	0.180 × 0.120 × 0.090	0.100 × 0.080 × 0.040
$hkl$ ranges	−13 ≤ $h$ ≤ 13 −18 ≤ $k$ ≤ 18 −18 ≤ $l$ ≤ 18	−15 ≤ $h$ ≤ 15 −24 ≤ $k$ ≤ 24 −24 ≤ $l$ ≤ 24
$\theta$ range (°)	1.886 to 34.487	2.076 to 30.569
Reflections collected/ unique/[Rint]	47884/8790/ [R <sub>int</sub> ] = 0.0343	96200/18771/ [R(int) = 0.01031]
Completeness to $\theta$ (%)	99.7	99.9
Absorption Correction	Semi-empirical from equivalents	Semi-empirical from equivalents
Max. and min. transmis.	0.7468 and 0.6880	0.7461 and 0.6818
Refinement method	Full matrix least-squares on $ F ^2$	Full-matrix least-squares on $ F ^2$
Data/restraints/parameters	8790/0/308	18771/1/902
Goodness of fit (GOF) on $ F ^2$	1.077	1.074
Final $R$ indices [ $I > 2\sigma(I)$ ]	$R_1 = 0.0384$ , $wR_2 = 0.1025$	$R_1 = 0.0628$ , $wR_2 = 0.1442$
$R$ indices (all data)	$R_1 = 0.0469$ , $wR_2 = 0.1102$	$R_1 = 0.1404$ , $wR_2 = 0.1947$
Extinction coefficient	n/a	n/a
Largest. Diff. Peak and hole (e Å <sup>-3</sup> )	1.540 and −0.544	2.970 and −1.091

Table 10. Crystal structure refinement parameters for 3 and 4.

	3	4
Empirical formula	C <sub>60</sub> H <sub>48</sub> Cd <sub>2</sub> N <sub>4</sub> O <sub>20</sub>	C <sub>113</sub> H <sub>92</sub> Cd <sub>4</sub> N <sub>16</sub> O <sub>41</sub>
Formula weight	1369.82	2779.62
$T$ (K)	100(2)	100(2)
Wavelength (Å)	0.71073	0.71073
System, space group	Triclinic, P-1	Monoclinic, P2/c
Unit cell dimensions		
$a$ (Å)	10.2552(6)	15.4163(10)
$b$ (Å)	11.9963(7)	17.3028(12)
$c$ (Å)	24.7728(15)	22.9004(17)
$\alpha$ (°)	82.798(2)	90
$\beta$ (°)	89.674(2)	105.565(2)
$\gamma$ (°)	65.1870(10)	90
$V$ (Å <sup>3</sup> )	2740.8(3)	5884.55
$Z$	2	2

Table 10. Cont.

	3	4
$D_{\text{calc}}$ (g cm <sup>3</sup> )	1.660	1.569
$\mu$ (mm <sup>-1</sup> )	0.862	0.806
$F$ (000)	1384	2804
Crystal size (mm <sup>3</sup> )	0.131 × 0.092 × 0.042	0.400 × 0.120 × 0.040
$hkl$ ranges	−13 ≤ $h$ ≤ 13 −15 ≤ $k$ ≤ 15 −32 ≤ $l$ ≤ 32	−22 ≤ $h$ ≤ 21 0 ≤ $k$ ≤ 25 0 ≤ $l$ ≤ 33
$\theta$ range (°)	2.191 to 27.526	2.190 to 31.067
Reflections collected/ unique/[Rint]	75664/12559/ [R <sub>int</sub> ] = 0.0750	18818/18818/ [R(int) = 0.0751]
Completeness to $\theta$ (%)	99.9	99.8
Absorption Correction	Semi-empirical from equivalents	Semi-empirical from equivalents
Max. and min. transmis.	0.7455 and 0.6770	0.7454 and 0.6631
Refinement method	Full matrix least-squares on $ F ^2$	Full-matrix least-squares on $ F ^2$
Data/restraints/parameters	12559/0/788	18818/10/776
Goodness of fit (GOF) on $ F ^2$	1.082	1.052
Final $R$ indices [ $I > 2\sigma(I)$ ]	$R_1 = 0.0404$ , $wR_2 = 0.0814$	$R_1 = 0.0428$ , $wR_2 = 0.1115$
$R$ indices (all data)	$R_1 = 0.0634$ , $wR_2 = 0.0888$	$R_1 = 0.0833$ , $wR_2 = 0.1439$
Extinction coefficient	n/a	n/a
Largest. Diff. Peak and hole (e Å <sup>-3</sup> )	1.087 and −1.028	2.986 and −1.223

**Methodology and computational details.** All of the calculations were performed using Gaussian09 software version D.01 [39]. The geometry optimization of the ground state and vertical absorptions from the electronically excited state for 1–4 were completed using density functional theory (DFT) and time-dependent DFT (TD-DFT), respectively, using  $\omega$ B97X-D [40,41] functional (S.I: Tables S3–S7 and Figures S28–S32). A correlation consistent basis set was used for the Zn, Cd, C, H, N and O atoms, the effective core potential CrenbL[42]. The MeOH solvation effects were incorporated using the polarizable continuum model-Linear Response (PCM-LR) [43,44]. Since either monomeric or dimeric arrays seem to be involved in absorption and emission depending on the concentration, the geometry of the monomer and the dimer in 2 were optimized separately. The frequencies were also computed for each optimized structure to ensure that the geometries corresponded to an energy minimum. The HOMO and LUMO energetic levels were firstly examined and then the energy gaps calculated. For 1–4, the first 80 vertical absorptions from the ground state to the excited states were calculated and only the most probable transitions, those with higher oscillator strength ( $f$ ) values, were selected for the electronic analysis. The shift in the theoretical absorption spectra, with respect to the experimental profile, is within the range of typical TD-DFT calculations (~0.3 eV) and are caused by computing the absorptions as vertical transitions [35].

The methodology for the generation of the MOs was the same as has been previously reported [15] (S.I: Figures S18–S22). The analysis of the electronic transitions was supported by the NTOs (S.I: Figures S23–S27) [45] to better identify and represent the main

contributor molecular orbitals of each transition. The NTOs were generated using Multiwfn software [46] version 3.7 with an isovalue of 0.02.

**Supplementary Materials:** The following are available online, calculated *S* values from the crystal structures of **1–4**, Table S1. Calculated *S* values from the optimized geometries of **1–4**, Table S2. FTIR-ATR, <sup>1</sup>H NMR, <sup>13</sup>C{<sup>1</sup>H} NMR and DEPT-135 spectra, Figures S1–S12. Solid state photoluminescence spectrum of **1**, Figure S13. Comparison between experimental and theoretical UV-Vis spectra of complexes **1**, **3** and **4**, Figures S14–S16. HOMO–LUMO energy gaps, Figure S17. MOs representation of **1–4**, Figures S18–S22. NTOs representation of **1–4**, Figures S23–S27. Details about DFT optimization of **1–4**, Tables S3–S7 and Figures S28–S32. Complete information about the crystal structure and molecular geometry is available in .cif format as Supporting Information. CCDC 2124892–2124895 contains the supplementary crystallographic data for this paper. These data can be retrieved free of charge via [www.ccdc.cam.ac.uk/data\\_request/cif](http://www.ccdc.cam.ac.uk/data_request/cif) (accessed on 15 January 2022), or by emailing [data\\_request@ccdc.cam.ac.uk](mailto:data_request@ccdc.cam.ac.uk), or by contacting The Cambridge Crystallographic Data Centre, 12 Union Road, Cambridge CB2 1EZ, UK; fax: +44-1223-336033.

**Author Contributions:** Conceptualization, J.P.; data curation, F.S.-F., J.M.R.-B. and M.F.-B.; formal analysis, F.S.-F. and M.F.-B.; funding acquisition, J.P.; investigation, F.S.-F.; methodology, F.S.-F. and J.M.R.-B.; project administration, J.P.; resources, T.C. and J.P.; software, F.S.-F. and J.M.R.-B.; supervision, J.P.; validation, J.A.A., T.C. and J.P.; visualization, F.S.-F. and J.M.R.-B.; writing—original draft, F.S.-F.; writing—review and editing, J.A.A., T.C. and J.P. All authors have read and agreed to the published version of the manuscript.

**Funding:** J.P. acknowledges financial support from the CB615921 project, the CB616406 project from “Fundació La Caixa” and the 2017SGR1687 project from the Generalitat de Catalunya.

**Informed Consent Statement:** Not applicable.

**Acknowledgments:** F.S.F. acknowledges the PIF pre-doctoral fellowship from the Universitat Autònoma de Barcelona.

**Conflicts of Interest:** The authors declare no conflict of interest.

**Sample Availability:** Samples of the compounds **1–4** are available from the authors.

## References

1. Gusev, A.; Shul, V.; Braga, E.; Zamnius, E.; Kryukova, M.; Linert, W. Luminescent Properties of Zn Complexes Based on Tetradentate N<sub>2</sub>O<sub>2</sub>-Donor Pyrazolone Schiff Bases. *Dyes Pigment.* **2020**, *183*, 108626. [CrossRef]
2. Lu, C.; Xu, Z.; Cui, J. Ratiometric and Highly Selective Fluorescent Sensor for Cadmium under Physiological PH Range: A New Strategy to Discriminate Cadmium from Zinc. *J. Org. Chem.* **2007**, *72*, 3554–3557. [CrossRef] [PubMed]
3. Chandra, A.; Das, M.; Pal, K.; Jana, S.; Dutta, B.; Ray, P.P.; Jana, K.; Sinha, C. Three-Dimensional-Coordination Polymer of Zn(II)-Carboxylate: Structural Elucidation, Photoelectrical Conductivity, and Biological Activity. *ACS Omega* **2019**, *4*, 17649–17661. [CrossRef] [PubMed]
4. Dutta, B.; Dey, A.; Sinha, C.; Pratim Ray, P.; Mir, M.H. Photochemical Structural Transformation of a Linear 1D Coordination Polymer Impacts the Electrical Conductivity. *Inorg. Chem.* **2018**, *57*, 8029–8032. [CrossRef]
5. Wu, S.; Zheng, Y.; Liu, Q.; Li, R.; Peng, T. Low Cost and Solution-Processable Zinc Phthalocyanine as Alternative Hole Transport Material for Perovskite Solar Cells. *RSC Adv.* **2016**, *6*, 107723–107731. [CrossRef]
6. Kumar, G.; Kumar, G.; Gupta, R. Effect of Pyridyl Donors from Organic Ligands: Versus Metalloligands on Material Design. *Inorg. Chem. Front.* **2021**, *8*, 1334–1373. [CrossRef]
7. Hancock, R.D. The Pyridyl Group in Ligand Design for Selective Metal Ion Complexation and Sensing. *Chem. Soc. Rev.* **2013**, *42*, 1500–1524. [CrossRef]
8. Barbieri, A.; Accorsi, G.; Armaroli, N. Luminescent Complexes beyond the Platinum Group: The D10 Avenue. *Chem. Commun.* **2008**, *19*, 2185–2193. [CrossRef]
9. Kokina, T.E.; Rakhmanova, M.I.; Shekhovtsov, N.A.; Glinskaya, L.A.; Komarov, V.Y.; Agafontsev, A.M.; Baranov, A.Y.; Plyusnin, P.E.; Sheludyakova, L.A.; Tkachev, A.V.; et al. Luminescent Zn(II) and Cd(II) Complexes with Chiral 2,2′-Bipyridine Ligands Bearing Natural Monoterpene Groups: Synthesis, Speciation in Solution and Photophysics. *Dalton Trans.* **2020**, *49*, 7552–7563. [CrossRef]
10. Dumur, F.; Contal, E.; Wantz, G.; Gimes, D. Photoluminescence of Zinc Complexes: Easily Tunable Optical Properties by Variation of the Bridge Between the Imido Groups of Schiff Base Ligands. *Eur. J. Inorg. Chem.* **2014**, 4186–4198. [CrossRef]
11. Vickers, S.M.; Frischmann, P.D.; Maclachlan, M.J. Family of Cadmium Acetate Coordination Networks With Structurally Diverse [Cd<sub>4</sub>(OAc)<sub>9</sub>(μ<sub>3</sub>-OH)]<sup>2-</sup> Secondary Building Units. *Inorg. Chem.* **2011**, *50*, 2957–2965. [CrossRef] [PubMed]

12. Cepeda, J.; Pérez-yáñez, S.; Rodríguez-diéguez, A.; García, J.Á.; Rojas, S. Towards Correlating Dimensionality and Topology in Luminescent MOFs Based on Terephthalato and Bispyridyl-like Ligands. *Dalton Trans.* **2021**, *50*, 9269–9282. [[CrossRef](#)] [[PubMed](#)]
13. Guerrero, M.; Vázquez, S.; Ayllón, J.A.; Calvet, T.; Font-Bardia, M.; Pons, J. Zn(II) and Cd(II) Coordination Dimers Based on Mixed Benzodioxole-Carboxylate and N-Donor Ligands: Synthesis, Characterization, Crystal Structures and Photoluminescence Properties. *ChemistrySelect* **2017**, *2*, 632–639. [[CrossRef](#)]
14. Moreno-Gómez, L.; Sánchez-Férez, F.; Calvet, T.; Font-Bardia, M.; Pons, J. Zn(II) and Cd(II) Monomer, Dimer and Polymer Compounds Coordinated by Benzoic Acid and 4-Acetylpyridine: Synthesis and Crystal Structures. *Inorg. Chim. Acta* **2020**, *506*, 119561. [[CrossRef](#)]
15. Sánchez-Férez, F.; Rius-Bartra, J.M.; Calvet, T.; Font-Bardia, M.; Pons, J. Steric and Electronic Effects on the Structure and Photophysical Properties of Hg(II) Complexes. *Inorg. Chem.* **2021**, *60*, 3851–3870. [[CrossRef](#)]
16. Deacon, G.B.; Phillips, R.J. Relationships between the Carbon-Oxygen Stretching Frequencies of Carboxylate Complexes and the Type of Carboxylate Coordination. *Coord. Chem. Rev.* **1980**, *33*, 227–250. [[CrossRef](#)]
17. Nandi, C.K.; Hazra, M.K.; Chakraborty, T. Vibrational Coupling in Carboxylic Acid Dimers. *J. Chem. Phys.* **2005**, *123*, 124310. [[CrossRef](#)]
18. Williams, D.H.; Fleming, I. *Spectroscopic Methods in Organic Chemistry*, 7th ed.; Springer Nature: Cham, Switzerland, 2008.
19. Nakamoto, K. *Infrared and Raman Spectra of Inorganic and Coordination Compounds Part A: Theory and Applications*, 6th ed.; Wiley: Hoboken, NJ, USA, 2009.
20. Pinsky, M.; Avnir, D. Continuous Symmetry Measures. 5. The Classical Polyhedra. *Inorg. Chem.* **1998**, *37*, 5575–5582. [[CrossRef](#)]
21. Addison, A.W.; Rao, T.N.; Reedijk, J.; van Rijn, J.; Verschoor, G.C. Synthesis, Structure, and Spectroscopic Properties of Copper(II) Compounds Containing Nitrogen-Sulphur Donor Ligands; the Crystal and Molecular Structure of Aqua[1,7-Bis(N-Methylbenzimidazol-2'-yl)-2,6-Dithiaheptane]Copper(II) Perchlorate. *J. Chem. Soc. Dalton Trans.* **1984**, 1349–1356. [[CrossRef](#)]
22. Nishio, M. The CH/ $\pi$  Hydrogen Bond in Chemistry. Conformation, Supramolecules, Optical Resolution and Interactions Involving Carbohydrates. *Phys. Chem. Chem. Phys.* **2011**, *13*, 13873–13900. [[CrossRef](#)]
23. Singh, J.; Hundal, G.; Gupta, R. Mononuclear and Dinuclear Ni<sup>II</sup> and Cu<sup>II</sup> Complexes with a Pyrrolicarboxamide Ligand: Core Conversions and Unusual Presence of a Dimer and Two Monomers in the Same Unit Cell. *Eur. J. Inorg. Chem.* **2009**, 3259–3265. [[CrossRef](#)]
24. Morse, P.M.; Girolami, G.S. Are d<sup>0</sup>ML<sub>6</sub> Complexes Always Octahedral? The X-Ray Structure of Trigonal-Prismatic [Li(Tmed)]<sub>2</sub>[ZrMe<sub>6</sub>]. *J. Am. Chem. Soc.* **1989**, *111*, 4114–4116. [[CrossRef](#)]
25. Friese, J.C.; Krol, A.; Puke, C.; Kirschbaum, K.; Giolando, D.M. Trigonal Prismatic vs Octahedral Coordination Geometry: Syntheses and Structural Characterization of Hexakis(Arylthiolato) Zirconate Complexes. *Inorg. Chem.* **2000**, *39*, 1496–1500. [[CrossRef](#)]
26. Llunell, M.; Casanova, D.; Cirera, J.; Bofill, J.M.; Alemany, P.; Alvarez, S.; Pinsky, M.; Avnir, D. *SHAPE Version 2.1: Program for the Stereochemical Analysis of Molecular Fragments by means of Continuous Shape Measures and Associated Tools*; Universitat de Barcelona: Barcelona, Spain, 2013.
27. Groom, C.R.; Bruno, I.J.; Lightfoot, M.P.; Ward, S.C. The Cambridge Structural Database. *Acta Crystallogr. Sect. B Struct. Sci. Cryst. Eng. Mater.* **2016**, *72*, 171–179. [[CrossRef](#)]
28. Shannon, R.D. Revised Effective Ionic Radii and Systematic Studies of Interatomic Distances in Halides and Chalcogenides. *Acta Crystallogr. Sect. A Found. Crystallogr.* **1976**, *32*, 751–767.
29. Alessi, P.J.; Carter, E.C.; Fairchild, M.D.; Hunt, R.W.G.; Kráncz, B.; Moore, J.R.; Morren, L.; Nobbs, J.H.; Ohno, Y.; Pointer, M.R.; et al. *CIE 15: Technical Report: Colorimetry*, 3rd ed.; Carter, E.C., Ohno, Y., Pointer, M.R., Robertson, A.R., Sève, R., Schanda, J.D., Witt, K., Eds.; International Commission on Illumination: Washington, DC, USA, 2004.
30. Würth, C.; Grabolle, M.; Pauli, J.; Spieles, M.; Resch-Genger, U. Relative and Absolute Determination of Fluorescence Quantum Yields of Transparent Samples. *Nat. Protoc.* **2013**, *8*, 1535–1550. [[CrossRef](#)]
31. Chen, R.F. Fluorescence Quantum Yields of Tryptophan and Tyrosine. *Anal. Lett.* **1967**, *1*, 35–42. [[CrossRef](#)]
32. Dawson, W.R.; Windsor, M.W. Fluorescence Yields of Aromatic Compounds. *J. Phys. Chem.* **1968**, *72*, 3251–3260. [[CrossRef](#)]
33. Hale, G.M.; Querry, M.R. Optical Constants of Water in the 200-Nm to 200-Mm Wavelength Region. *Appl. Opt.* **1973**, *12*, 555. [[CrossRef](#)]
34. Shakerizadeh-Shirazi, F.; Hemmateenejad, B.; Mehranpour, A.M. Determination of the Empirical Solvent Polarity Parameter ET(30) by Multivariate Image Analysis. *Anal. Methods* **2013**, *5*, 891–896. [[CrossRef](#)]
35. Adamo, C.; Jacquemin, D. The Calculations of Excited-State Properties with Time-Dependent Density Functional Theory. *Chem. Soc. Rev.* **2013**, *42*, 845–856. [[CrossRef](#)] [[PubMed](#)]
36. Sheldrick, G.M. A Short History of SHELX. *Acta Crystallogr. Sect. A Found. Crystallogr.* **2008**, *64*, 112–122. [[CrossRef](#)]
37. MacRae, C.F.; Sovago, I.; Cottrell, S.J.; Galek, P.T.A.; McCabe, P.; Pidcock, E.; Platings, M.; Shields, G.P.; Stevens, J.S.; Towler, M.; et al. Mercury 4.0: From Visualization to Analysis, Design and Prediction. *J. Appl. Crystallogr.* **2020**, *53*, 226–235. [[CrossRef](#)] [[PubMed](#)]
38. Persistence of Vision Pty. Ltd. *Persistence of Vision*; Persistence of Vision Pty. Ltd.: Williamstown, VIC, Australia, 2004.
39. Frisch, M.J.; Trucks, G.W.; Schlegel, H.B.; Scuseria, G.E.; Robb, M.A.; Cheeseman, J.R.; Scalmani, G.; Barone, V.; Mennucci, B.; Petersson, G.A.; et al. *Gaussian 09, version D.01*; Gaussian, Inc.: Wallingford, UK, 2016.

40. Chai, J.D.; Head-Gordon, M. Systematic Optimization of Long-Range Corrected Hybrid Density Functionals. *J. Chem. Phys.* **2008**, *128*, 084106. [[CrossRef](#)] [[PubMed](#)]
41. Chai, J.D.; Head-Gordon, M. Long-Range Corrected Hybrid Density Functionals with Damped Atom-Atom Dispersion Corrections. *Phys. Chem. Chem. Phys.* **2008**, *10*, 6615–6620. [[CrossRef](#)] [[PubMed](#)]
42. LaJohn, L.A.; Christiansen, P.A.; Ross, R.B.; Atashroo, T.; Ermler, W.C. Ab Initio Relativistic Effective Potentials with Spin-Orbit Operators. III. Rb through Xe. *J. Chem. Phys.* **1987**, *87*, 2812–2824. [[CrossRef](#)]
43. Improta, R.; Barone, V.; Scalmani, G.; Frisch, M.J. A State-Specific Polarizable Continuum Model Time Dependent Density Functional Theory Method for Excited State Calculations in Solution. *J. Chem. Phys.* **2006**, *125*, 054103. [[CrossRef](#)]
44. Improta, R.; Scalmani, G.; Frisch, M.J.; Barone, V. Toward Effective and Reliable Fluorescence Energies in Solution by a New State Specific Polarizable Continuum Model Time Dependent Density Functional Theory Approach. *J. Chem. Phys.* **2007**, *127*, 074504. [[CrossRef](#)]
45. Martin, R.L. Natural Transition Orbitals. *J. Chem. Phys.* **2003**, *118*, 4775–4777. [[CrossRef](#)]
46. Lu, T.; Chen, F. Multiwfn: A Multifunctional Wavefunction Analyzer. *J. Comput. Chem.* **2012**, *33*, 580–592. [[CrossRef](#)]