

# Luminescence of Divalent Ytterbium in $\text{Ca}_2\text{SiS}_4$ .

Anthony B. Parmentier<sup>1,\*</sup>, Philippe F. Smet<sup>1</sup>, and Dirk Poelman<sup>1</sup>

<sup>1</sup> Lumilab, Department of Solid State Sciences, Ghent University, Krijgslaan 281/S1, B-9000 Ghent, Belgium

\* Corresponding author: Anthony.Parmentier@UGent.be

## Abstract

The luminescence of ytterbium doped calcium thiosilicate is studied. Excitation and emission spectra are presented and related to an energy level scheme. A comparison with divalent europium in the same host is carried out.

## 1. Introduction

Calcium thiosilicate is one of the most interesting host for divalent europium within the class of sulfides [1]. It provides two different  $\text{Ca}^{2+}$ -sites for which the  $\text{Eu}^{2+}$  can substitute. The  $\text{Eu}^{2+}$  emission spectrum contains two maxima, one peaking at 560nm and one at 660nm, and each peak can be related to a crystallographic site [2]. From the well documented trends in the behaviour of different rare earth ions doped in the same host material, interesting luminescent properties are expected to occur for  $\text{Yb}^{2+}$  doped  $\text{Ca}_2\text{SiS}_4$  [3].

## 2. Experimental details

$\text{Ca}_2\text{SiS}_4$  was prepared by solid state reaction and doped with Yb (concentrations at%: 0.1% and 1%), as well as with Eu, Sm and Tm for comparison, as these are the divalent ions for which 5d-4f luminescence is most likely to be observed. XRD was used for checking phase composition. Photoluminescence excitation and emission spectra and thermal quenching profiles were obtained using a helium cryostat in a steady state photoluminescence spectrometer (Edinburgh Instruments FS 920).

## 3. Results and discussion

For Sm and Tm, the emission spectrum is dominated by 4f-4f emission from the trivalent state of these ions.

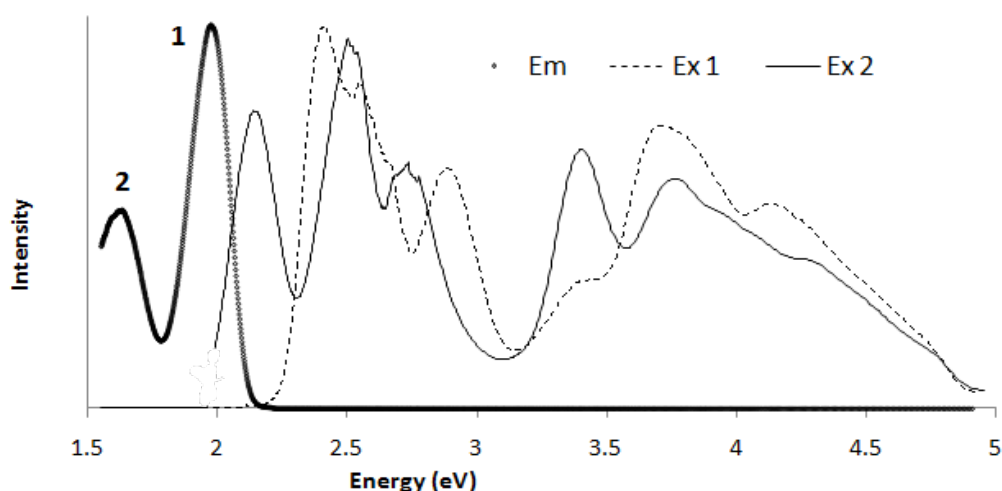


Figure 1: Excitation (Ex1 monitored at peak1, Ex2 at peak 2) and emission spectrum (at 70K) of  $\text{Ca}_2\text{SiS}_4:\text{Yb}^{2+}$

Ytterbium in  $\text{Ca}_2\text{SiS}_4$  does exhibit 5d-4f luminescence, but this emission is seriously quenched at room temperature. Therefore, all spectra presented in this work are recorded at low temperature.

### 3.1. Emission and excitation spectra

The emission spectrum of  $\text{Ca}_2\text{SiS}_4:\text{Yb}^{2+}$  consists of two peaks, as is the case for  $\text{Eu}^{2+}$  in the same host. In **Figure 1**, the emission peak with the shortest wavelength (1.97eV; 630nm) is labeled “1”. The one with the longer wavelength (1.62eV; 765nm) is labeled “2”. The emission peaks for  $\text{Eu}^{2+}$  in the same host are found at 2.21eV (560nm) and 1.88eV (660nm). In contrast with the case of  $\text{Ca}_2\text{SiS}_4:\text{Eu}^{2+}$ , the excitation spectrum of  $\text{Ca}_2\text{SiS}_4:\text{Yb}^{2+}$  reveals clear and resolved substructure. From these data, an energy level scheme for  $\text{Yb}^{2+}$  in this material is proposed.

### 3.2. Temperature dependent behaviour

The thermal quenching temperature (**Figure 2**) of  $\text{Yb}^{2+}$  in calcium thiosilicate is around 230K, considerably lower than 470K for  $\text{Eu}^{2+}$  in the same host. This is attributed to the lower activation energy for  $\text{Yb}^{2+}$  to excite the 5d electron to conduction band states. [4,5].

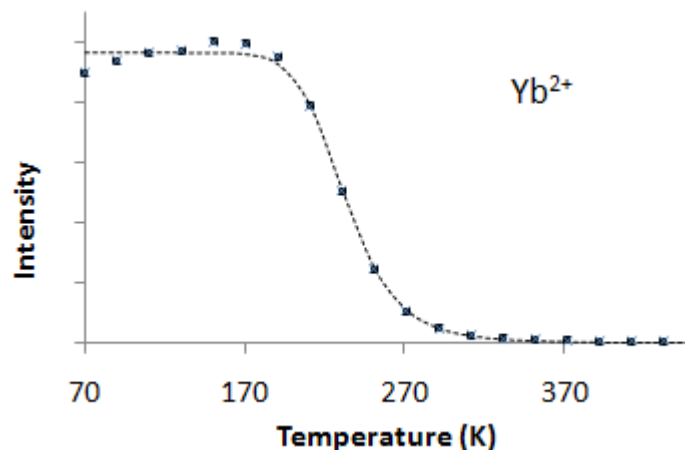


Figure 2. Thermal quenching behaviour of  $\text{Yb}^{2+}$  in  $\text{Ca}_2\text{SiS}_4$ .

Both emission peaks have a distinctly different quenching behaviour, which is analysed in detail and compared with the quenching of  $\text{Eu}^{2+}$ -luminescence in the same material.

### 3.3. Decay time measurements

Decay time measurements are performed. The decay profile at 70K can be approximated by a double exponential, in which the shortest wavelength peak (1 in Figure 1, 1.97eV; 630nm) has the longest lifetime and the longest wavelength peak (2 in Figure 1, 1.62eV; 765nm) has the shortest lifetime.

## References

- [1] P.F. Smet, I. Moreels, Z. Hens, D. Poelman, *Materials* 3 (2010) 2834-2883.
- [2] P. F. Smet, N. Avci, B. Loos, J. E. Van Haecke and D. Poelman, *J. Phys.: Condens. Matter* 19 (2007) 246223
- [3] P. Dorenbos, *J. Phys.: Condens. Matter* 15 (2003) 575–594
- [4] S. Lizzo, E. P. Klein Nagelvoort, R. Erens, A. Meijerink and G. Blasse, *J. Phys. Chem Solids* 58 (1997) 963-968
- [5] P. Dorenbos, *J. Phys.: Condens. Matter* 15 (2003) 2645–2665