

## UvA-DARE (Digital Academic Repository)

### Interconversion of Phosphinyl Radical and Phosphinidene Complexes by Proton Coupled Electron Transfer

Abbenseth, J.; Delony, D.; Neben, M.C.; Würtele, C.; de Bruin, B.; Schneider, S.

**DOI**

[10.1002/anie.201901470](https://doi.org/10.1002/anie.201901470)

[10.1002/ange.201901470](https://doi.org/10.1002/ange.201901470)

**Publication date**

2019

**Document Version**

Final published version

**Published in**

Angewandte Chemie, International Edition

**License**

CC BY-NC-ND

[Link to publication](#)

**Citation for published version (APA):**

Abbenseth, J., Delony, D., Neben, M. C., Würtele, C., de Bruin, B., & Schneider, S. (2019). Interconversion of Phosphinyl Radical and Phosphinidene Complexes by Proton Coupled Electron Transfer. *Angewandte Chemie, International Edition*, 58(19), 6338-6341. <https://doi.org/10.1002/anie.201901470>, <https://doi.org/10.1002/ange.201901470>

**General rights**

It is not permitted to download or to forward/distribute the text or part of it without the consent of the author(s) and/or copyright holder(s), other than for strictly personal, individual use, unless the work is under an open content license (like Creative Commons).

**Disclaimer/Complaints regulations**

If you believe that digital publication of certain material infringes any of your rights or (privacy) interests, please let the Library know, stating your reasons. In case of a legitimate complaint, the Library will make the material inaccessible and/or remove it from the website. Please Ask the Library: <https://uba.uva.nl/en/contact>, or a letter to: Library of the University of Amsterdam, Secretariat, Singel 425, 1012 WP Amsterdam, The Netherlands. You will be contacted as soon as possible.

UvA-DARE is a service provided by the library of the University of Amsterdam (<https://dare.uva.nl>)

## Proton Coupled Electron Transfer

International Edition: DOI: 10.1002/anie.201901470  
German Edition: DOI: 10.1002/ange.201901470

## Interconversion of Phosphinyl Radical and Phosphinidene Complexes by Proton Coupled Electron Transfer

Josh Abbenseth, Daniel Delony, Marc C. Neben, Christian Würtele, Bas de Bruin,\* and Sven Schneider\*

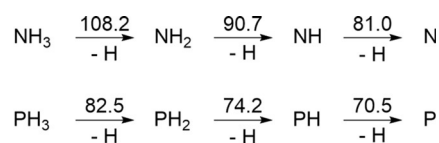
This paper is dedicated to Professor Thomas Fässler on the occasion of his 60th birthday

**Abstract:** The isolable complex  $[\text{Os}(\text{PHMes}^*)\text{H}(\text{PNP})]$  ( $\text{Mes}^* = 2,4,6\text{-}i\text{-Bu}_3\text{C}_6\text{H}_3$ ;  $\text{PNP} = \text{N}\{\text{CHCHP}^i\text{Bu}_2\}_2$ ) exhibits high phosphinyl radical character. This compound offers access to the phosphinidene complex  $[\text{Os}(\text{PMes}^*)\text{H}(\text{PNP})]$  by P–H proton coupled electron transfer (PCET). The P–H bond dissociation energy (BDE) was determined by isothermal titration calorimetry and supporting DFT computations. The phosphinidene product exhibits electrophilic reactivity as demonstrated by intramolecular C–H activation.

Electronically unsaturated phosphorous compounds, such as phosphinyl radicals ( $\text{PR}_2$ ), are key transient species in P–C bond forming reactions, like olefin phosphination.<sup>[1,2]</sup> Structural and spectroscopic characterization of free and coordinated phosphinyl radicals facilitated the examination of (electronic) structure/reactivity relationships.<sup>[3,4]</sup> Free phosphinidenes (PR) could only very recently be sufficiently stabilized.<sup>[5]</sup> While a variety of transition metals form isolable phosphinidene complexes ( $\text{M} = \text{PR}$ ) with promising stoichiometric reactivity, such as P–C bond formation,<sup>[6,7]</sup> catalytic phosphinidene transfer protocols remain rare,<sup>[8]</sup> compared to, for example, nitrene transfer which has emerged as a powerful method for C–N bond formation.<sup>[9]</sup>

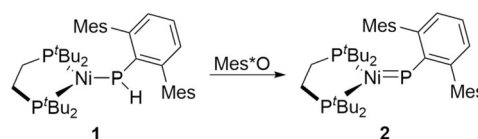
The scarcity of catalytic PR-transfer might in part be attributed to a lack of suitable oxidizing PR-transfer reagents, such as analogs of iminoiodane ( $\text{ArI}=\text{NR}$ ) or azide ( $\text{RN}_3$ )

nitrene sources. Primary phosphines are attractive precursors from a thermochemical point of view due to lower intrinsic P–H compared to N–H bond dissociation (free) energies (BD(F)Es; Scheme 1).<sup>[10,11]</sup> The P–H bonds might be further



**Scheme 1.** N–H and P–H bond dissociation energies (BDEs) in the gas phase ( $\text{kcal mol}^{-1}$ ).<sup>[10]</sup>

weakened upon coordination to a metal catalyst as was demonstrated, for example, for coordinated  $\text{H}_2\text{O}$  and  $\text{NH}_3$  ligands.<sup>[12]</sup> Amide ligand N–H bond activation by proton coupled electron transfer (PCET) is an active field of research.<sup>[13]</sup> In contrast, phosphide P–H homolysis was only reported by Hillhouse and co-workers (Scheme 2).<sup>[14]</sup> Inspired by this precedence, we examined P–H PCET as an entry to phosphinidene chemistry. Herein, the formation of a rare phosphinyl radical complex and its transformation to a terminal osmium phosphinidene complex by PCET are reported.



**Scheme 2.** Generation of a nickel phosphinidene complex by PCET reported by Hillhouse and co-workers. ( $\text{Mes} = 2,4,6\text{-Me}_3\text{C}_6\text{H}_3$ ,  $\text{Mes}^* = 2,4,6\text{-}i\text{-Bu}_3\text{C}_6\text{H}_3$ ).<sup>[14b]</sup>

We previously utilized the four-coordinate complex  $[\text{Os}^{\text{II}}\text{Cl}(\text{PNP})]$  (**3**,  $\text{PNP} = \text{N}\{\text{CHCHP}^i\text{Bu}_2\}_2$ ; Scheme 3) to stabilize low-valent imide and nitride complexes.<sup>[15]</sup> Complex **3** readily reacts with  $\text{PH}_2\text{Mes}^*$  in  $\text{Et}_2\text{O}$  to the dark blue P–H oxidative addition product  $[\text{OsCl}(\text{PHMes}^*)\text{H}(\text{PNP})]$  (**4**) (Scheme 3). Complex **4** decomposes over the course of a few hours in solution at room temperature but could be fully characterized including single crystal X-ray diffraction.<sup>[16]</sup> NMR spectra are in agreement with  $C_1$  symmetry at  $-30^\circ\text{C}$ . All signals ( $^1\text{H}$ ,  $^{13}\text{C}$ ,  $^{31}\text{P}$ ) are broadened supporting hindered rotation around the Os–P bond. The phosphide P–

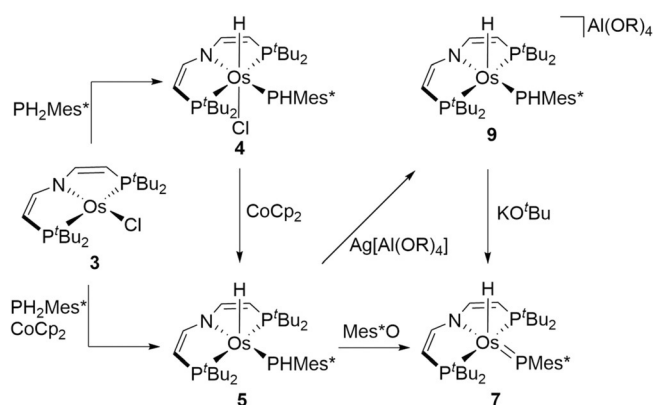
[\*] M. Sc. J. Abbenseth, M. Sc. D. Delony, B. Sc. M. C. Neben, Dr. C. Würtele, Prof. Dr. S. Schneider  
Georg-August-Universität Göttingen, Institut für Anorganische Chemie

Tammannstraße 4, 37077 Göttingen (Germany)  
E-mail: sven.schneider@chemie.uni-goettingen.de

Prof. Dr. B. de Bruin  
Van't Hoff Institute for Molecular Sciences (HIMS), University of Amsterdam (UvA)  
Science Park 904, 1098 XH Amsterdam (The Netherlands)  
E-mail: b.debruin@uva.nl

Supporting information and the ORCID identification number(s) for the author(s) of this article can be found under:  
<https://doi.org/10.1002/anie.201901470>.

© 2019 The Authors. Published by Wiley-VCH Verlag GmbH & Co. KGaA. This is an open access article under the terms of the Creative Commons Attribution Non-Commercial NoDerivs License, which permits use and distribution in any medium, provided the original work is properly cited, the use is non-commercial, and no modifications or adaptations are made.



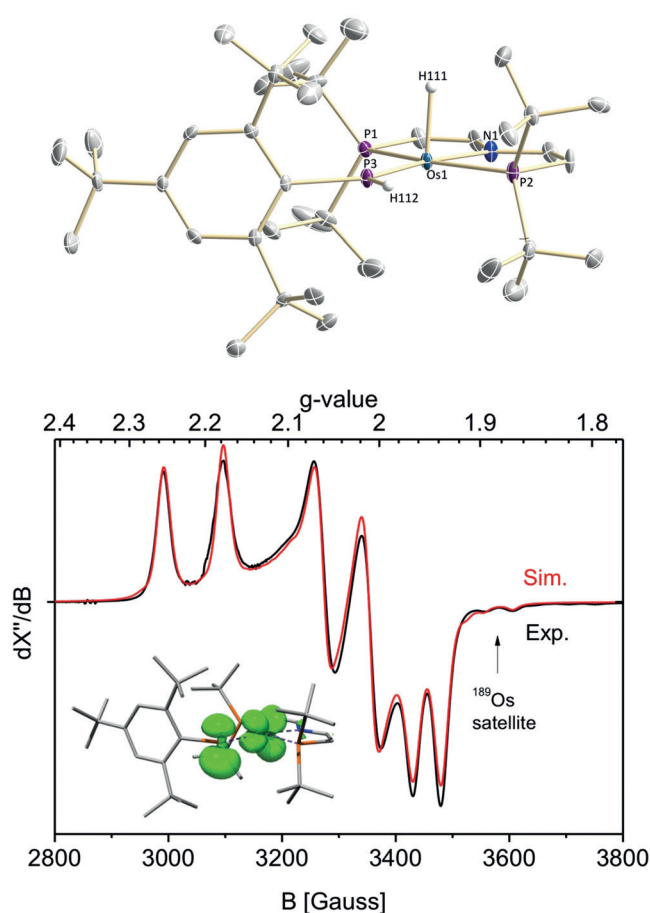
**Scheme 3.** Synthetic access to phosphanide and phosphinidene complexes **4–6** and **9** ( $R = C(CF_3)_3$ ).

atom exhibits trigonal-planar coordination in the solid state, indicating P to  $Os^{IV}$   $\pi$  donation.<sup>[16]</sup>

Reduction of **4** with cobaltocene affords the radical product  $[OsH(PHMes^*)(PNP)]$  (**5**) in 80% isolated yield (Scheme 3). The magnetic moment in solution ( $\mu_{eff} = 1.51 \mu_B$ ) derived by Evan's method is in agreement with an  $S = 1/2$  system with unquenched orbital momentum. Complex **5** can also be directly synthesized from **3** by one pot P–H oxidative addition and subsequent reduction. Characteristic bands for the  $PHMes^*$  and hydride ligands were detected in the IR spectrum ( $\nu_{PH} = 2345 \text{ cm}^{-1}$ ;  $\nu_{OsH} = 2180 \text{ cm}^{-1}$ ). In the solid state, complex **5** features distorted square pyramidal coordination geometry with the hydride ligand in the apical position (Figure 1). Upon reduction from **4** to **5** the Os–P bond to the  $PHMes^*$  ligand is slightly elongated by  $\Delta d = 0.04 \text{ \AA}$ .

The EPR spectrum of **5** in frozen toluene exhibits a rhombic signal that is in agreement with an  $S = 1/2$  low-spin configuration (Figure 1). The  $g$ -anisotropy is considerably smaller than that of typical osmium(III) complexes, in line with reduced spin-orbit interaction due to ligand redox non-innocence.<sup>[15a,17]</sup> Accordingly, large and slightly rhombic hyperfine interaction (HFI) with one  $^{31}P$  nucleus is observed in all principal directions of the  $g$ -tensor (Figure 1), supporting considerable spin delocalization to the  $PHMes^*$  ligand. The free phosphinyl radical  $PPh_2$  exhibits axial  $^{31}P$ -HFI with an isotropic coupling constant ( $A_{iso}(^{31}P) = 260 \text{ MHz}$ ) close to that of **5** ( $A_{iso}(^{31}P) = 201 \text{ MHz}$ ).<sup>[18]</sup> The higher isotropic HFI, yet reduced dipolar coupling, found for the transient phosphanyl radical complex  $[W(PPh_2)(CO)_5]$  ( $A_{iso}(^{31}P) = 499 \text{ MHz}$ ) was attributed to  $p_P \rightarrow d_W$  spin delocalization ( $\rho_P = 75\%$ ) and concomitant polarization of the P lone-pair.<sup>[4a]</sup> For **5**, the higher  $g$ -anisotropy and smaller isotropic and anisotropic contributions of the  $^{31}P$ -HFI tensor are consistent with increased P  $\rightarrow$  M spin delocalization. Comparing the isotropic and anisotropic ( $T_x = -144 \text{ MHz}$ ,  $T_y = -96 \text{ MHz}$ ,  $T_z = +239 \text{ MHz}$ ) contributions to the  $^{31}P$ -HFI with atomic parameters allows for a rough estimate of phosphorous spin densities ( $\rho_{3s} \approx 2\%$ ;  $\rho_{3p} \approx 42\%$ ) when treating the HFI as approximately axial ( $T_{\perp} = (T_x + T_y)/2$ ).<sup>[18]</sup>

DFT calculations fully support this interpretation. Both the molecular structure and the EPR parameters (Figure 1) were excellently reproduced computationally. The SOMO of



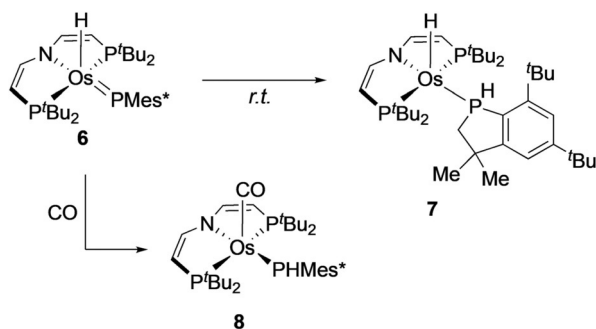
**Figure 1.** Top: Molecular structure of **5** in the solid state from single-crystal X-ray diffraction (thermal ellipsoids set at 50% probability); solvent molecules, anions, and hydrogen atoms, except the P–H and Os–H hydrogen atoms, are omitted for clarity. Selected bond lengths [ $\text{\AA}$ ] and angles [ $^\circ$ ]: **5** Os1–H111 1.69(6), Os1–N1 2.069(4), Os1–P1 2.3723(12), Os1–P2 2.3537(13), Os1–P3 2.2301(13), P3–H112 1.33(5); P1–Os1–P2 160.20(4); N1–Os1–P3 172.17(11). Bottom: X-band EPR spectrum of **5** (9.4366 GHz) in toluene at 148 K (black) and simulation (red) with the following parameters (DFT computed values in brackets):  $g_x = 1.952$  (1.951),  $g_y = 2.033$  (2.024),  $g_z = 2.214$  (2.230);  $A_x(^{31}P) = +57 \text{ MHz}$  (+58 MHz),  $A_y(^{31}P) = +105 \text{ MHz}$  (+95 MHz),  $A_z(^{31}P) = +440 \text{ MHz}$  (+441 MHz); Euler angles:  $\alpha = -73$  ( $-74.3$ ),  $\beta = +137$  (+142.7),  $\gamma = -7.5$  ( $-7.5$ );  $A_x(^{189}Os) = -240$  ( $-277$ ) MHz. Insert: Computed spin density plot.

**5** represents an antisymmetric ( $\pi^*$ ) combination of the metal  $d_{yz}$  orbital and a phosphorus  $p$ -orbital. Reduced  $\pi$ -bonding is expressed by the Os– $PHMes^*$  Mayer bond index (1.46). In consequence, the computed spin density (Figure 1) is almost equally distributed between the P ( $\rho_P = 47\%$ ) and Os atoms ( $\rho_{Os} = 50\%$ ).

Complex **5** represents an unprecedented, isolable phosphanide complex with large phosphinyl redox non-innocent character ( $Os^{II}-PR_2$ ). We therefore examined PCET reactivity of the radical ligand, specifically as an entry to phosphinidene chemistry. Two experimental methods are widely used to estimate E–H BD(F)Es, that is, a) bracketing based on hydrogen transfer reactions with reference H-donor/acceptor reagents and b) quantification by the “square-

Scheme" formalism, that is, a thermochemical redox/protonation cycle.<sup>[19]</sup>

Complex **5** shows no reactivity with the H-atom donor TEMPO-H (TEMPO = 2,2,6,6-tetramethylpiperidinyloxy;  $BDFE_{O-H} = ca. 66 \text{ kcal mol}^{-1}$ ),<sup>[19]</sup> indicating weak P–H bonds for the hypothetical osmium(II) phosphine complex [Os(PH<sub>2</sub>Mes\*)H(PNP)].<sup>[20]</sup> In turn, **5** readily reacted with the H-atom acceptors Mes\*O and TEMPO. A purple product is obtained in yields around 90% with low thermal stability at room temperature even in the solid state (see below) but could be characterized by NMR spectroscopy at  $-30^\circ\text{C}$ . Retention of the hydride ligand is indicated by a <sup>1</sup>H NMR signal at  $\delta_H = -15.9$  ppm. Formation of the phosphinidene complex [Os(PMes\*)H(PNP)] (**6**) is evidenced by the <sup>31</sup>P NMR signal at  $\delta_P = 825$  ppm, that is, assignable to the PMes\* group. While suitable crystals for X-ray diffraction could not be obtained, the structural integrity of the osmium pincer phosphinidene framework is supported by the spectroscopic characterization of the C–H insertion product **7** as a mixture of two diastereomers (Scheme 4). Such electrophilic phosphinidene insertion has previously been reported.<sup>[6d,21]</sup> Furthermore, addition of CO to **6** gives the five-coordinate phosphinidene complex **8** (Scheme 4) after Os–H reductive elimination.<sup>[16]</sup>

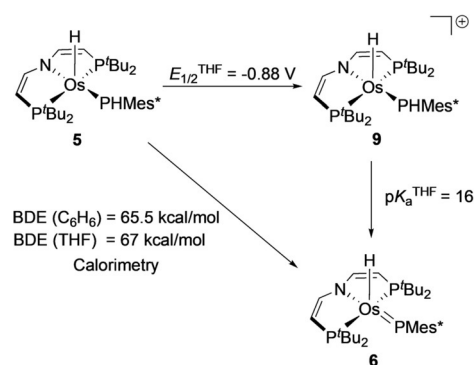


**Scheme 4.** Reactivity of phosphinidene complex **6**.

Parent **5** offers two potential sites for H-atom transfer (HAT). The generation of **6** indicates higher Os–H over P–H bond strength if the reaction proceeds under thermodynamic control. BDE quantification was attempted by stepwise oxidation and deprotonation. The cyclic voltammogram of **5** reveals quasi-reversible reduction at  $E_{p,c} = -2.06 \text{ V}$  and reversible oxidation  $E_{1/2} = -0.88 \text{ V}$  (vs.  $\text{FcCp}_2^{+/0}$ ).<sup>[16]</sup> Chemical oxidation with  $\text{Ag}[\text{Al}(\text{OC}(\text{CF}_3)_3)_4]$  at  $-35^\circ\text{C}$  immediately gives the deep blue osmium(IV) phosphide complex **9** (Scheme 3). Complex **9** readily decomposes at room temperature but could be characterized at low temperatures including crystallography.<sup>[16]</sup> Deprotonation of in situ prepared **9** with  $\text{KO}t\text{Bu}$  at  $-80^\circ\text{C}$  gives phosphinidene **6** almost quantitatively. However, the low thermal stability of **9** hampered reliable  $pK_a$  determination.

The P–H bond strength of **5** was therefore derived by isothermal titration calorimetry (ITC). Titration of **5** with Mes\*O in benzene or THF afforded the reaction enthalpies for HAT ( $\Delta_r H^{C_6H_6} = -16.5 \text{ kcal mol}^{-1}$ ,  $\Delta_r H^{THF} = -17 \text{ kcal mol}^{-1}$ ) and consequently the  $BDE_{P-H}$  of **5** ( $BDE^{C_6H_6} =$

$65.1 \pm 1 \text{ kcal mol}^{-1}$ ,  $BDE^{THF} = 67 \pm 1 \text{ kcal mol}^{-1}$ ).<sup>[22]</sup> DFT analysis of **5** supports the experimental P–H BDE (calibrated  $BDE_{P-H} = 67.5 \text{ kcal mol}^{-1}$ ; non-calibrated value:  $64.0 \text{ kcal mol}^{-1}$ ),<sup>[23]</sup> that is, considerably lower than the Os–H bond (calibrated  $BDE_{P-H} = 74.2 \text{ kcal mol}^{-1}$ ; non-calibrated:  $70.1 \text{ kcal mol}^{-1}$ ), indicating that phosphinidene **6** is the thermodynamic PCET product. The calorimetric and electrochemical data also allows for calculating the  $pK_a$  of **9** ( $pK_a^{THF} = 16$ ) from a thermochemical square-scheme (Scheme 5).<sup>[24]</sup>



**Scheme 5.** Thermochemical square-scheme of **5**, **6** and **9**.

In conclusion, we presented the first spectroscopically and crystallographically characterized phosphide complex with large phosphanyl radical character. The rhombic <sup>31</sup>P–HFI tensor and the DFT model are in line with even spin delocalization over the Os–P core. Versatile access to an electrophilic phosphinidene complex that undergoes intramolecular C–H activation was demonstrated by P–H PCET. Thermochemical analysis by means of ITC was utilized due to thermal instability of **9**. The data indicates that concerted or stepwise ET/PT are both viable routes from phosphide to phosphinidene complexes.

## Acknowledgements

This work was funded by the European Research Council (ERC Grant Agreement 646747), the Fond der Chemischen Industrie (FCI Doktoranden Stipendium for J.A.), the Deutsche Forschungsgemeinschaft (DFG, 389479699/GRK2455) and The Netherlands Organisation for Scientific Research (NWO TOP-Grant 716.015.001) for financial support. Furthermore, the authors thank Dr. A. C. Stückl and R. Schöne for EPR and NMR measurements, respectively.

## Conflict of interest

The authors declare no conflict of interest.

**Keywords:** osmium · phosphanyl radical · phosphinidene · pincer · proton coupled electron transfer

**How to cite:** *Angew. Chem. Int. Ed.* **2019**, *58*, 6338–6341  
*Angew. Chem.* **2019**, *131*, 6404–6407

- [1] S. Marque, P. Tordo, *Top. Curr. Chem.* **2005**, *250*, 43.
- [2] a) P. P. Power, *Chem. Rev.* **2003**, *103*, 789; b) A. Armstrong, T. Chivers, R. T. Boere, *ACS Symp. Ser.* **2006**, *917*, 66; c) C. D. Martin, M. Soleilhavoup, G. Bertrand, *Chem. Sci.* **2013**, *4*, 3020; d) V. Nesterov, D. Reiter, P. Bag, P. Frisch, R. Holzner, A. Porzelt, S. Inoue, *Chem. Rev.* **2018**, *118*, 9678.
- [3] a) S. L. Hinchley, C. A. Morrison, D. W. H. Rankin, C. L. B. Macdonald, R. J. Wiacek, A. Voigt, A. H. Cowley, M. F. Lappert, G. Gundersen, J. A. C. Clyburne, P. P. Power, *J. Am. Chem. Soc.* **2001**, *123*, 9045; b) J. P. Bezombes, K. B. Borisenko, P. B. Hitchcock, M. F. Lappert, J. E. Nycz, D. W. H. Rankin, H. E. Robertson, *Dalton Trans.* **2004**, 1980; c) S. Ito, M. Kikuchi, M. Yoshifuji, A. J. Arduengo III, T. A. Konovalova, L. D. Kispert, *Angew. Chem. Int. Ed.* **2006**, *45*, 4341; *Angew. Chem.* **2006**, *118*, 4447; d) R. Kinjo, B. Donnadiou, G. Bertrand, *Angew. Chem. Int. Ed.* **2010**, *49*, 5930; *Angew. Chem.* **2010**, *122*, 6066; e) O. Back, M. A. Celik, G. Frenking, M. Melaimi, B. Donnadiou, G. Bertrand, *J. Am. Chem. Soc.* **2010**, *132*, 10262; f) O. Back, B. Donnadiou, M. von Hopffgarten, S. Klein, R. Tonner, G. Frenking, G. Bertrand, *Chem. Sci.* **2011**, *2*, 858; g) S. Ishida, F. Hirakawa, T. Iwamoto, *J. Am. Chem. Soc.* **2011**, *133*, 12968; h) X. Pan, X. Wang, Y. Zhao, Y. Sui, X. Wang, *J. Am. Chem. Soc.* **2014**, *136*, 9834; i) K. Schwedtmann, S. Schulz, F. Hennersdorf, T. Strassner, E. Dmitrieva, J. J. Weigand, *Angew. Chem. Int. Ed.* **2015**, *54*, 11054; *Angew. Chem.* **2015**, *127*, 11206; j) K. Schwedtmann, G. Zanon, J. J. Weigand, *Chem. Asian J.* **2018**, *13*, 1388.
- [4] a) B. Ndiaye, S. Bhat, A. Jouaiti, T. Berclaz, G. Bernardinelli, M. Geoffroy, *J. Chem. Phys. A* **2006**, *110*, 9736; b) P. Agarwal, N. A. Piro, K. Meyer, P. Muller, C. C. Cummins, *Angew. Chem. Int. Ed.* **2007**, *46*, 3111; *Angew. Chem.* **2007**, *119*, 3171; c) M. Scheer, C. Kuntz, M. Stubenhofer, M. Linseis, R. F. Winter, M. Sierka, *Angew. Chem. Int. Ed.* **2009**, *48*, 2600; *Angew. Chem.* **2009**, *121*, 2638; d) G. Tan, J. Li, L. Zhang, C. Chen, Y. Zhao, X. Wang, Y. Song, Y.-Q. Zhang, M. Driess, *Angew. Chem. Int. Ed.* **2017**, *56*, 12741; *Angew. Chem.* **2017**, *129*, 12915; e) U. Fischbach, M. Trincado, H. Grützmacher, *Dalton Trans.* **2017**, *46*, 3443; f) Y.-E. Kim, Y. Lee, *Angew. Chem. Int. Ed.* **2018**, *57*, 14159; *Angew. Chem.* **2018**, *130*, 14355.
- [5] a) L. Liu, D. A. Ruiz, D. Munz, G. Bertrand, *Chem* **2016**, *1*, 147; b) M. M. Hansmann, R. Jassar, G. Bertrand, *J. Am. Chem. Soc.* **2016**, *138*, 8356.
- [6] a) K. Lammertsma, *Top. Curr. Chem.* **2003**, *229*, 95; b) J. C. Slootweg, K. Lammertsma in *Science of Synthesis*, Vol. 42 (Eds.: B. M. Trost, F. Mathey), Thieme, Stuttgart, **2009**, p. 15; c) R. Waterman, *Dalton Trans.* **2009**, 18; d) H. Aktaş, J. C. Slootweg, K. Lammertsma, *Angew. Chem. Int. Ed.* **2010**, *49*, 2102; *Angew. Chem.* **2010**, *122*, 2148.
- [7] M. J. Amme, A. B. Kazi, T. R. Cundari, *Int. J. Quantum Chem.* **2010**, *110*, 1702.
- [8] a) K. Pal, O. B. Hemming, B. M. Day, T. Pugh, D. J. Evans, R. A. Layfield, *Angew. Chem. Int. Ed.* **2016**, *55*, 1690; *Angew. Chem.* **2016**, *128*, 1722; b) J. K. Pagano, B. J. Ackley, R. Waterman, *Chem. Eur. J.* **2018**, *24*, 2554.
- [9] a) G. Dequierez, V. Pons, P. Dauban, *Angew. Chem. Int. Ed.* **2012**, *51*, 7384; *Angew. Chem.* **2012**, *124*, 7498; b) J. L. Roizen, M. E. Harvey, J. Du Bois, *Acc. Chem. Res.* **2012**, *45*, 911; c) P. F. Kuijpers, J. I. van der Vlugt, S. Schneider, B. de Bruin, *Chem. Eur. J.* **2017**, *23*, 13819.
- [10] a) N. V. Barkovskii, V. I. Tsirel'nikov, A. M. Emel'yanov, Y. S. Khodoev, *Teplofiz. Vys. Temp.* **1991**, *29*, 474; b) K. P. Huber, G. Herzberg, *Molecular Spectra and Molecular Structure of Diatomic Molecules*, Van Nostrand, New York, **1979**, p. 456; c) J. Berkowitz, L. A. Curtiss, S. T. Gibson, J. P. Greene, G. L. Hillhouse, J. A. Pople, *J. Chem. Phys.* **1986**, *84*, 375; d) J. D. Cox, D. D. Wagman, V. A. Medvedev, *CODATA Key Values for Thermodynamics*, Hemisphere Publishing Corp., New York, **1984**, p. 1; e) "NIST-JANAF Thermochemical Tables, Fourth Edition": M. W. Chase, Jr., *J. Phys. Chem. Ref. Data Monogr.* **1998**, *9*, 1.
- [11] The BDE(NH<sub>2</sub>) was calculated using a Hess cycle: BDE(NH<sub>2</sub>) = Δ<sub>f</sub>H<sub>gas</sub><sup>o</sup>(N) + 2Δ<sub>f</sub>H<sub>gas</sub><sup>o</sup>(H) – BDE(NH) – Δ<sub>f</sub>H<sub>gas</sub><sup>o</sup>(NH<sub>2</sub>).
- [12] a) S. S. Kolmar, J. M. Mayer, *J. Am. Chem. Soc.* **2017**, *139*, 10687; b) M. J. Bezdek, S. Guo, P. J. Chirik, *Science* **2016**, *354*, 730.
- [13] a) D. J. Mindiola, G. L. Hillhouse, *J. Am. Chem. Soc.* **2001**, *123*, 4623; b) M. H. Huynh, T. J. Meyer, *Proc. Natl. Acad. Sci. USA* **2004**, *101*, 13138; c) R. E. Cowley, R. P. Bontchev, J. Sorrell, O. Sarracino, Y. Feng, H. Wang, J. M. Smith, *J. Am. Chem. Soc.* **2007**, *129*, 2424; d) I. Nieto, F. Ding, R. P. Bontchev, H. Wang, J. M. Smith, *J. Am. Chem. Soc.* **2008**, *130*, 2716; e) R. E. Cowley, N. A. Eckert, S. Vaddadi, T. M. Figg, T. R. Cundari, P. L. Holland, *J. Am. Chem. Soc.* **2011**, *133*, 9796; f) V. M. Iluc, A. J. M. Miller, J. S. Anderson, M. J. Monreal, M. P. Mehn, G. L. Hillhouse, *J. Am. Chem. Soc.* **2011**, *133*, 13055; g) S. Wiese, J. L. McAfee, D. R. Pahls, C. L. McMullin, T. R. Cundari, T. H. Warren, *J. Am. Chem. Soc.* **2012**, *134*, 10114; h) C. Milsmann, S. P. Semproni, P. J. Chirik, *J. Am. Chem. Soc.* **2014**, *136*, 12099; i) M. G. Scheibel, J. Abbenseth, M. Kinauer, F. W. Heinemann, C. Würtele, B. de Bruin, S. Schneider, *Inorg. Chem.* **2015**, *54*, 9290; j) I. Pappas, P. J. Chirik, *J. Am. Chem. Soc.* **2016**, *138*, 13379; k) D. M. Spasyuk, S. H. Carpenter, C. E. Kefalidis, W. E. Piers, M. L. Neidig, L. Maron, *Chem. Sci.* **2016**, *7*, 5939; l) M. J. Bezdek, P. J. Chirik, *Angew. Chem. Int. Ed.* **2018**, *57*, 2224; *Angew. Chem.* **2018**, *130*, 2246.
- [14] a) R. Melenkivitz, D. J. Mindiola, G. L. Hillhouse, *J. Am. Chem. Soc.* **2002**, *124*, 3846; b) V. M. Iluc, G. L. Hillhouse, *J. Am. Chem. Soc.* **2010**, *132*, 15148.
- [15] a) J. Abbenseth, M. Diefenbach, S. C. Bete, C. Würtele, C. Volkmann, S. Demeshko, M. C. Holthausen, S. Schneider, *Chem. Commun.* **2017**, *53*, 5511; b) J. Abbenseth, S. C. Bete, M. Finger, C. Volkmann, C. Würtele, S. Schneider, *Organometallics* **2018**, *37*, 802.
- [16] See the Supporting Information for spectroscopic, electrochemical, crystallographic or computational details.
- [17] J. A. McCleverty, T. J. Meyer in *Comprehensive Coordination Chemistry II*, Vol. 5, Elsevier, Amsterdam, **2003**, p. 556.
- [18] M. Geoffroy, E. A. C. Lucken, C. Mazeline, *Mol. Phys.* **1974**, *28*, 839.
- [19] J. J. Warren, T. A. Tronic, J. M. Mayer, *Chem. Rev.* **2010**, *110*, 6961.
- [20] Steric reasons are unlikely for the lack of reactivity as **5** readily undergoes P–H transfer with TEMPO.
- [21] J. D. Masuda, K. C. Jantunen, O. V. Ozerov, K. J. T. Noonan, D. P. Gates, B. L. Scott, J. L. Kiplinger, *J. Am. Chem. Soc.* **2008**, *130*, 2408.
- [22] The BDE<sub>O–H</sub> of Mes\*OH in THF was calculated using the pK<sub>a</sub> value from (see the Supporting Information): F. Ding, J. M. Smith, H. Wang, *J. Org. Chem.* **2009**, *74*, 2679.
- [23] The calculated BDEs were calibrated (taking advantage of internal error cancellation) against the experimentally available BDE of propene, using a Hess cycle (see Supporting Information).
- [24] E. P. Cappellani, S. D. Drouin, G. Jia, P. A. Maltby, R. H. Morris, C. T. Schweitzer, *J. Am. Chem. Soc.* **1994**, *116*, 3375.

Manuscript received: February 1, 2019

Accepted manuscript online: March 6, 2019

Version of record online: April 1, 2019