

Research Article

Halloysite Nanotubes Noncovalently Functionalised with SDS Anionic Surfactant and PS-*b*-P4VP Block Copolymer for Their Effective Dispersion in Polystyrene as UV-Blocking Nanocomposite Films

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A simple and versatile method is reported for the noncovalent functionalisation of natural and “green” halloysite nanotubes (HNTs) allowing their effective dispersion in a polystyrene (PS) thermoplastic matrix via solvent mixing. Initially, HNTs (pristine HNTs) were modified with physically adsorbed surfactant molecules of sodium dodecyl sulphate (SDS) and PS-*b*-P4VP [P4VP: poly(4-vinylpyridine)] block copolymer (BCP). Hereafter, SDS and BCP modified HNTs will be indicated as SDS-*m*-HNT and BCP-*m*-HNT. Nanocomposite films with 1, 2, and 5 wt.% HNT loadings were prepared, abbreviated as PS-SDS-*m*-HNT1, PS-SDS-*m*-HNT2, and PS-SDS-*m*-HNT5 and PS-BCP-*m*-HNT1, PS-BCP-*m*-HNT2, and PS-BCP-*m*-HNT5 (where 1, 2, and 5 correspond to the wt.% of HNTs). All nanocomposites depicted improved thermal degradation compared to the neat PS as revealed by thermogravimetric analysis (TGA). Transmission electron microscopy (TEM) confirmed the good dispersion state of HNTs and the importance of modification by SDS and BCP. X-ray diffraction (XRD) studies showed the characteristic interlayer spacing between the two silicate layers of pristine and modified HNTs. The PS/HNT nanocomposite films exhibited excellent ultraviolet-visible (UV-vis) absorbance properties and their potential application as UV-filters could be envisaged.

1. Introduction

In the past decades, polymer nanocomposites reinforced with different kinds of nanoscale materials such as natural mineral layered silicate clays, carbon nanotubes (CNTs), and graphite nanoplatelets have attracted a steadily growing scientific interest due to their peculiar and fascinating properties, as well as their unique applications in commercial sectors [1–9]. The properties of the resulting nanocomposites depend greatly on the physical and chemical properties of the polymer matrices, the nature and surface chemistry of nanofillers,

and the method of preparation. In all cases, the uniform and homogeneous dispersion of nanofillers in the polymer matrixes, overcoming the tendency of nanoparticles to form agglomerates, is a general prerequisite for achieving desired mechanical and physical characteristics [10–13].

Naturally occurring clay-based polymer nanocomposites have attracted considerable attention owing to their significantly improved mechanical and thermal properties compared to conventional composites, even at low clay loadings [14–20]. The clay minerals investigated are principally lamellar clays and most notably laponite [21, 22],

smectite [23], kaolinite [24], and montmorillonite [25]. A natural clay material named as halloysite [26, 27] and its application in composites are still significantly unexplored. HNTs are naturally occurring two-layered aluminosilicate [$\text{Al}_2(\text{OH})_4\text{Si}_2\text{O}_5 \cdot 2\text{H}_2\text{O}$] with a predominantly hollow tubular structure, with sizes ranging in the submicron range, chemically similar to kaolin, mined from natural deposits in countries such as America, China, New Zealand, France, and Belgium [28]. Mismatch in the two-layered alignment of the tetrahedral sheet of silica bonded to the octahedral or gibbsite sheet of alumina causes the wall to curve into the cylindrical shape as explained by Bates et al. [29]. Intercalated water may occur between the repetitive two-layered sheets comprising the spiral wall, which tends to be irreversibly removed on drying [28]. Halloysite is a low cost alternative green material that can be mined from the consequent deposit as a raw mineral. The size of halloysite particles varies between 1 and 15 μm in length and 10 and 150 nm in diameter, depending on the deposits. Due to their high aspect ratio, HNTs give a large amount of filler-polymer interactions compared to other nanofillers. Moreover, the large diameter of the halloysite lumen is expected to accommodate some polymer molecules, which can offer polymeric composites with some specific properties like high degradation temperature and so forth [28]. Therefore, HNTs have been used as a new type of filler for polymers, that is, epoxies, polypropylene, polyvinyl alcohol, and so on, to improve the mechanical and thermal properties of the composites [30–33]. It has been shown also that polymer/HNT nanocomposites exhibited significantly reduced gas permeability, increased thermal stability, and improved flame retardancy [28, 30, 34, 35].

Similar to other clay materials, HNTs exhibit a large amount of surface hydroxyl groups rendering their dispersion into a polymer matrix and mostly to nonpolar ones very difficult [28, 36]. This phenomenon is very familiar with well-established silica filler. Bifunctional organosilanes, such as bis(triethoxysilylpropyl)-tetrasulfide (TESPT), have been proposed to chemically modify silica surfaces for that purpose [30]. Yuan et al. [37] reported the modification of HNT by amino silane (3-aminopropyl) trimethoxysilane (APTES) and tried to understand the reaction chemistry behind it. Silane bearing glycidyl group treated HNTs were incorporated to a cyanate ester cured epoxy resin and have been found to enhance the dispersion of HNTs [38]. Most of the work published until now is related to the functionalisation of HNTs and it is limited to the mechanical reinforcement of polymer matrices. The noncovalent functionalisation of HNTs and their dispersion in a PS thermoplastic matrix for the production of nanocomposites with exceptional UV absorbance and blocking properties are yet to be explored.

Herein, a facile approach is presented for the noncovalent modification of HNTs and the preparation of PS/HNT nanocomposites with improved dispersion and exceptional UV absorbance properties. The improvement in the degree of dispersion of HNTs in the PS matrix is the main objective of the present work because the dispersion of HNTs in polymers is a difficult issue due to their hydrophilic nature. The first modification is achieved using the SDS anionic surfactant

and the second using a BCP of the polystyrene-block-poly-4-vinylpyridine (PS-b-P4VP) type. The outer surface of HNTs was modified by physical adsorption of SDS and BCP molecules, which gave a stable suspension of HNTs in PS solutions and a good dispersion in the final PS nanocomposite films after evaporation of the solvent. XRD was used to study the crystal structure of pristine HNTs. Fourier-transform infrared spectroscopy (FT-IR) and TGA prove the successful modification of HNTs with SDS and BCP molecules. The microstructure of pristine and modified HNTs was investigated with field-emission scanning electron microscopy (FE-SEM) and high resolution transmission electron microscopy (HR-TEM). TGA and DSC were used to study the structural decomposition and the crystallisation temperature of PS/HNT nanocomposites. Thin sections of PS/HNT nanocomposites were prepared at room temperature via ultramicrotomy and TEM micrographs were taken showing the effect of the noncovalent modification on the dispersion of HNTs in the PS matrix. The variation of absorbance and transmittance of the PS/HNT nanocomposite films with respect to neat PS film in the UV and visible range of light was studied by UV-vis spectroscopy. The excellent UV absorbance blocking properties of the nanocomposite films suggest their potential use as UV-filters.

2. Experimental

2.1. Materials. Polystyrene (PS) with average molecular weight (M_w) of $M_w \sim 400,000$ was purchased from Sigma Aldrich and used as the matrix for the fabrication of PS/HNT nanocomposite films. Halloysite nanotubes (aspect ratio: $L/D = 4$, according to the supplier's data) and sodium dodecyl sulphate (SDS) were obtained from OmniPur. PS-b-P4VP (product ID: P4829P-S4VP) diblock copolymer with the following molecular characteristics, $M_w/M_n = 1.05$, $M_n(\times 10^3) = 50.0\text{-b-}13.0$, was obtained from Polymer Source. Tetrahydrofuran (THF), the solvent for preparing the PS solutions and the respective PS/HNT nanocomposites, was obtained from Sigma Aldrich. All chemicals were of analytical grade and used as received without further purification.

2.2. Modification of HNTs with SDS and BCP and Fabrication of PS/HNT Nanocomposite Films. In order to modify HNTs with the anionic surfactant SDS, 0.02 wt.% of HNTs were mixed with 1 wt.% of SDS aqueous solution, stirred for 24 h, and sonicated for 5 h. The mixture was kept to stand for 24 h and the supernatant HNTs were collected, centrifuged, and dried in a vacuum oven at 70°C for 24 h. For modification of HNTs by BCP, a solution of PS-b-P4VP with (1 wt.%) was prepared in THF and the same amount of HNTs was added and sonicated for 3 h. The solutions were kept to stand for 24 h, and the supernatant HNTs were collected, centrifuged, and dried in a vacuum oven at 50°C for 24 h.

For the PS nanocomposites, initially homopolymer PS was dissolved in THF, 10% (w/v). Then, 1, 2, and 5 wt.% of SDS-m-HNT and BCP-m-HNT were added to the PS solutions in THF (wt.% with respect to the polymer mass in the solution). These solutions were sonicated for 10 min and then magnetically stirred for 1 h. Finally, the PS/HNT solutions

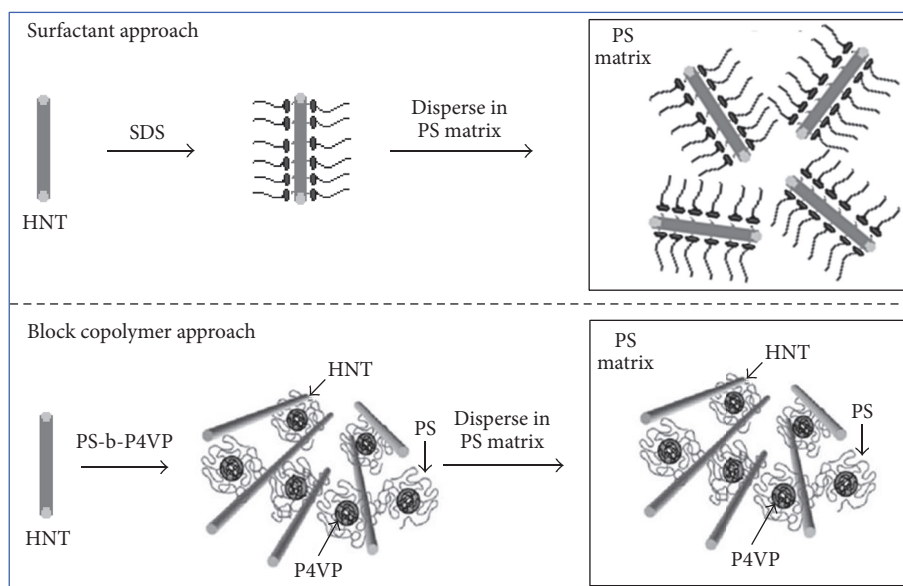


FIGURE 1: Schematic diagram showing the mechanism for the physical adsorption of SDS and BCP molecules onto the surface of HNTs and their effective dispersion in PS matrix.

were casted in Teflon moulds and dried at room temperature for 48 h. Finally, all the nanocomposite films were inserted in a vacuum oven and kept at 50°C under vacuum overnight to fully remove traces of solvent and induce a thermal annealing effect to the nanocomposite films for the better thermodynamic organisation of the polymer chains.

Figure 1 demonstrates the mechanism for the physical adsorption (noncovalent functionalisation) of SDS and BCP molecules to the surface of HNTs and the affinity of the organic pending groups towards the effective dispersion of HNTs in the PS matrix.

2.3. Characterisation Techniques. XRD was performed with an X-ray diffractometer XRD T/T (GE Inspection Technologies Ahrensburg, Germany) in symmetric step-scan mode with $\Delta 2\theta = 0.05^\circ$ in transmission operating at 40 kV and 30 mA with Cu K α radiation. FT-IR spectra were recorded using a Vertex 80v FT-IR spectrometer (Bruker Germany) equipped with a DTGS detector by signal averaging of 256 scans. Approximately 1.0 mg of pristine HNTs, SDS-m-HNT, and BCP-m-HNT was pressed together with 100 mg of crystalline KBr to form pellets. The pellets were scanned over the range 4000–400 cm^{-1} at a resolution of $\pm 2 \text{ cm}^{-1}$. Thermal analysis experiments were carried out using DSC and TGA. TGA studies were performed using a TA instrument (Q 5000) for the HNTs and PS/HNT nanocomposites. The temperature range varied from ambient (25°C) up to 800°C (for samples of approx. 10 mg), with 2 K/min heating rate, under nitrogen (N_2) atmosphere. Curie point standards were utilized for the temperature calibration. The initial degradation temperatures and the temperatures corresponding to 50% degradation of the samples were recorded. DSC measurements were carried out using 5–10 mg of PS/HNT nanocomposite samples to determine the glass transition temperatures (T_g) of the films. For the analysis, a DSC Q1000 from TA Instruments was used

at a 10 K/min heating rate under nitrogen atmosphere. The samples were heated to 140°C, held for 5 minutes for removing moisture and previous thermal history, and then cooled to room temperature at a cooling rate of 10°C/min. The second heating scans were used for determination of glass transition temperature of the samples. FE-SEM was performed using the NEON 40 (Carl Zeiss AG, Germany) scanning electron microscope. All the images were generated using the secondary electron detector (SE2) under an accelerating voltage of 1.0 kV. Samples were prepared by drop-casting 100 μL of pristine and modified HNTs on a $2 \times 1 \text{ cm}^2$ silicon substrate followed by drying at room temperature in a fume hood. TEM investigations were performed to study the geometry of HNTs and the microstructures of PS/HNT nanocomposite films using the Libra 200 transmission electron microscope (Carl Zeiss AG, Germany) operating at 200 kV. Samples for TEM were prepared by dispensing 10 μL of pristine HNTs on a Cu grid with a carbon support membrane followed by drying. Thin sections of approximately 70 nm thickness were obtained by ultramicrotomy using a Leica Ultramicrotome. Reichert Ultracut S. Diamond knives (Diatome) were used for the trimming (model cryotrim 45°) as well as the thin-section cutting process (model cryo 35°) and all preparation steps were done at room temperature as described elsewhere [39–41]. The thin sections were placed on copper grids (300 mesh Cu, Agar) and sputtered with a 10 nm thickness carbon layer to avoid charging effects and protect the samples from the electron beam [42]. The variation of absorbance and transmittance of the PS/HNT nanocomposite films containing 1, 2, and 5 wt.% SDS and BCP modified HNTs with respect to pure PS film in the UV and visible range of light was studied at room temperature using a Perkin-Elmer Lambda 800 UV-vis spectrophotometer (with an incorporated xenon flash lamp). The variation of transmittance of nanocomposites films with the amount of HNTs is also studied and discussed in detail.

3. Results and Discussion

3.1. Characterisation of HNTs

3.1.1. X-Ray Diffraction Analysis of Pristine and Modified HNTs. Figure 2 shows the XRD pattern of pristine HNTs, as well as SDS-m-HNT and BCP-m-HNT. Indexing can be done based on the presence of two different crystalline phases, namely, the halloysite and the kaolinite $[KAl_3(SO_4)_2(OH)_6]$ crystal structure. The close structural relationship between them gives rise to similar powder diffraction patterns. Over the years, this subject has been a big issue in the literature and nowadays it is generally accepted that, in the case of halloysite, the intensity ratio between the 002 and 020/110 reflections should be $\sim \leq 2$, which is true in case of our pristine HNT sample. The pure HNTs show a diffraction peak at $2\theta = 12.25$, which is related to the (0 0 1) plane. This basal reflection of HNTs is coming from its tubular morphology, high degree of disorder, small crystal size, and interstratifications of layer with various hydration states. This peak of pristine HNTs is shifted to a lower 2θ value for the SDS-m-HNT and BCP-m-HNT, respectively. The d -spacing of SDS-m-HNT is 0.73 nm at a 2θ of 12.15. The value of 2θ is even lower for the BCP-m-HNT sample, which is 11.85 corresponding to the 0.745 nm basal spacing. The observed reduction of 2θ corresponding to increased basal spacing of HNTs may be attributed to the intercalation of HNTs by SDS surfactant molecules and BCP micelles physical adsorption, respectively. Apart from the aforementioned peak at low 2θ , the diffraction pattern of pristine HNT is showing other peaks at 2θ of 20.06 and 24.94, respectively, which are related to (0 2 0) and (0 0 2) basal reflection. It should be noted that the peak at 20.06 almost vanishes in the SDS-m-HNT and BCP-m-HNT, while the peak at 24.94 is shifted to lower 2θ value supporting the intercalation of SDS molecules and BCP chains into the HNT gallery.

3.1.2. FT-IR Spectra of Pristine and Modified HNTs. FT-IR spectra of pristine, SDS, and BCP modified HNTs are shown in Figure 3. In FT-IR spectra of pristine HNTs, there is a very strong absorption at 1000 cm^{-1} due to the O-Si-O siloxane surface groups confirming of HNTs. Also the Al-OH peaks at 3694 , 3626 , and 3623 cm^{-1} and Si-O vibrations at 1119 and 1000 cm^{-1} were observed, which are characteristic peaks of HNTs. The following peaks are also seen: 906 cm^{-1} corresponding to O-H deformation of inner hydroxyl groups, 796 cm^{-1} corresponding to symmetric stretching of Si-O, and 684 cm^{-1} corresponding to perpendicular Si-O stretching. In the case of SDS-m-HNT, a peak corresponding to SO_3^- stretch at 1217 and 1248 cm^{-1} , CH_2 asymmetric stretch at 2917 , CH_2 symmetric stretch at 2850 , and CH_2 rocking mode at 721 cm^{-1} could be observed apart from the characteristic peaks of HNTs. The additional characteristic peaks of pyridine ring of PS-P4VP at 1416 and 1598 cm^{-1} are present in the BCP modified HNT sample. Figure 3(b) shows the colloidal stability of SDS and BCP modified HNTs well dispersed in the PS solution where no precipitation has been observed within 24 h. In case of SDS modification, the anionic surfactant

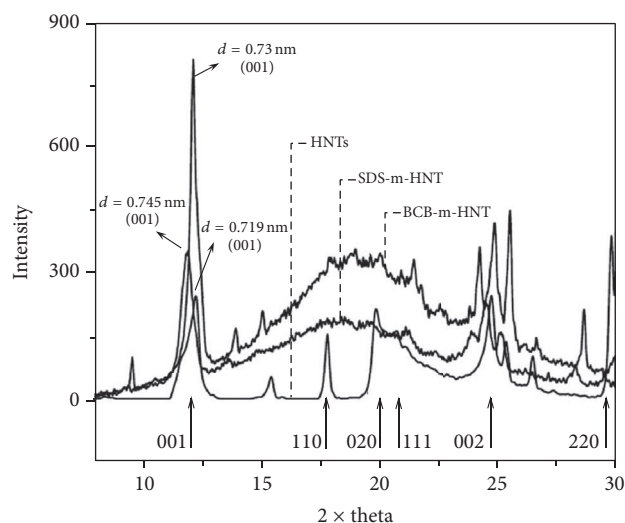


FIGURE 2: XRD pattern of pristine and modified HNTs.

forms a bilayer on the surface of HNTs making it organophilic and dispersible in the polymer matrix. The PS-*b*-P4VP as an amphiphilic block copolymer forms micelles, which act as dispersants by physically adhering to the surface of the HNTs.

3.1.3. TGA Analysis of Pristine, BCP, and SDS Modified HNTs.

Figure 4(a) shows the TGA and Figure 4(b) the derivative of weight loss for pristine, BCP, and SDS modified HNTs demonstrating their structural decomposition process and as a means to prove the noncovalent attachment of SDS and BCP molecules. As can be seen, the decomposition of pristine HNTs takes place in two main steps. At first, the endothermic peak around 100°C is due to the moisture present in the sample (1.6 wt.%). Structural decomposition of halloysite and alunite takes place in the second endothermic peak from 400 to 550°C (15.8 wt.%), and a residue of 82.8 wt.% was found at 650°C . In Figure 4(b), also the decomposition peaks of pristine HNTs (near to 100°C) corresponding to evaporation of moisture and the other at 450°C corresponding to structural decomposition of HNTs are observed. SDS modified HNTs show first weight loss at 200°C corresponding to SDS (27.5 wt.%), while decomposition of halloysite and alunite takes place in the second endothermic peak from 400 to 550°C (16.2 wt.%) with a final residue of 44.5 wt.% at 650°C . From Figure 4(b) DTG graph, for SDS modified HNTs first peak near 200°C corresponding to decomposition of SDS and the second at 420°C corresponding to decomposition of HNT can be seen. BCP modified HNTs show weight loss between 200 and 450°C with a final residue of 44.6 wt.% at 650°C . Considering the decomposition of halloysite and alunite from 400 to 550°C (16.0–16.5 wt.%), it is estimated that the amount of BCP chains noncovalently bound onto the surface of HNTs is approximately 37 wt.%. From Figure 4(b) for BCP modified HNTs, there is one large peak at around 400°C due to overlap of decomposition temperatures of PS-*b*-P4VP BCP and HNTs.

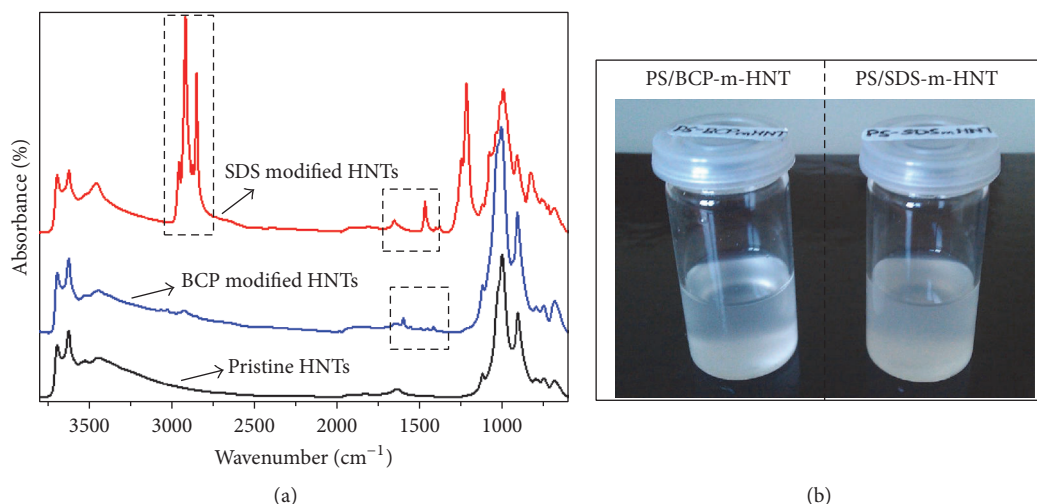


FIGURE 3: (a) Infrared spectra of pristine, BCP, and SDS modified HNTs; (b) PS solution containing BCP and SDS modified HNTs.

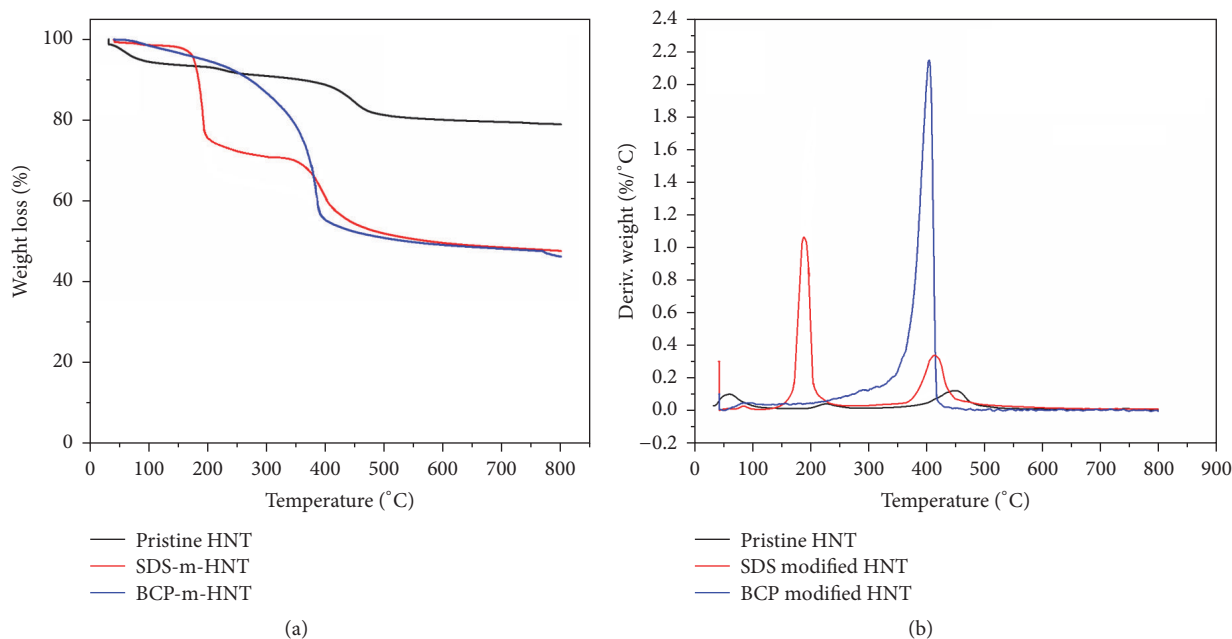


FIGURE 4: (a) TGA and (b) derivative of weight loss (DTG) of pristine, SDS, and BCP modified HNTs.

3.1.4. Scanning and Transmission Electron Microscopy of HNTs. Figure 5 shows the SEM images of pristine (Figures 5(a) and 5(b)), SDS modified (Figure 5(c)), and BCP modified (Figure 5(d)) HNTs, respectively. The diameters of pristine HNTs are found out to be in the range of 70–100 nm and the lengths in the range of 180–1300 nm. In Figure 5(d), it is observed that HNTs are wrapped by the BCPs molecules. Figures 6(a) and 6(b) are TEM images at two different magnifications depicting the geometrical characteristics of HNTs (HNT length and diameter), which as can be seen are in good agreement with the SEM investigations.

3.2. Characterisation of PS-HNT Nanocomposite Films. PS-HNT nanocomposite films were subjected to TGA analysis

(under N_2 atmosphere) in order to analyse the effect of HNT loading on the thermal stability and potential improvement. The results of TGA and DTG of PS-HNT nanocomposites containing SDS modified HNTs are shown in Figures 7(a) and 7(b), respectively. It is observed that the temperature required for 10% degradation, termed as the degradation onset temperature ($T_{0.1}$), is found to be increased from 250°C for pure PS film to 262°C, 276°C, and 289°C for nanocomposite films containing 1, 2, and 5% SDS modified HNTs, respectively (values given in Table 1). Also, the temperature required for 50% degradation ($T_{0.5}$) is increased from 385°C for pure PS film to 394°C, 400°C, and 410°C for nanocomposite films containing 1, 2, and 5% SDS modified HNTs. It can be concluded that the thermal stability of PS nanocomposite

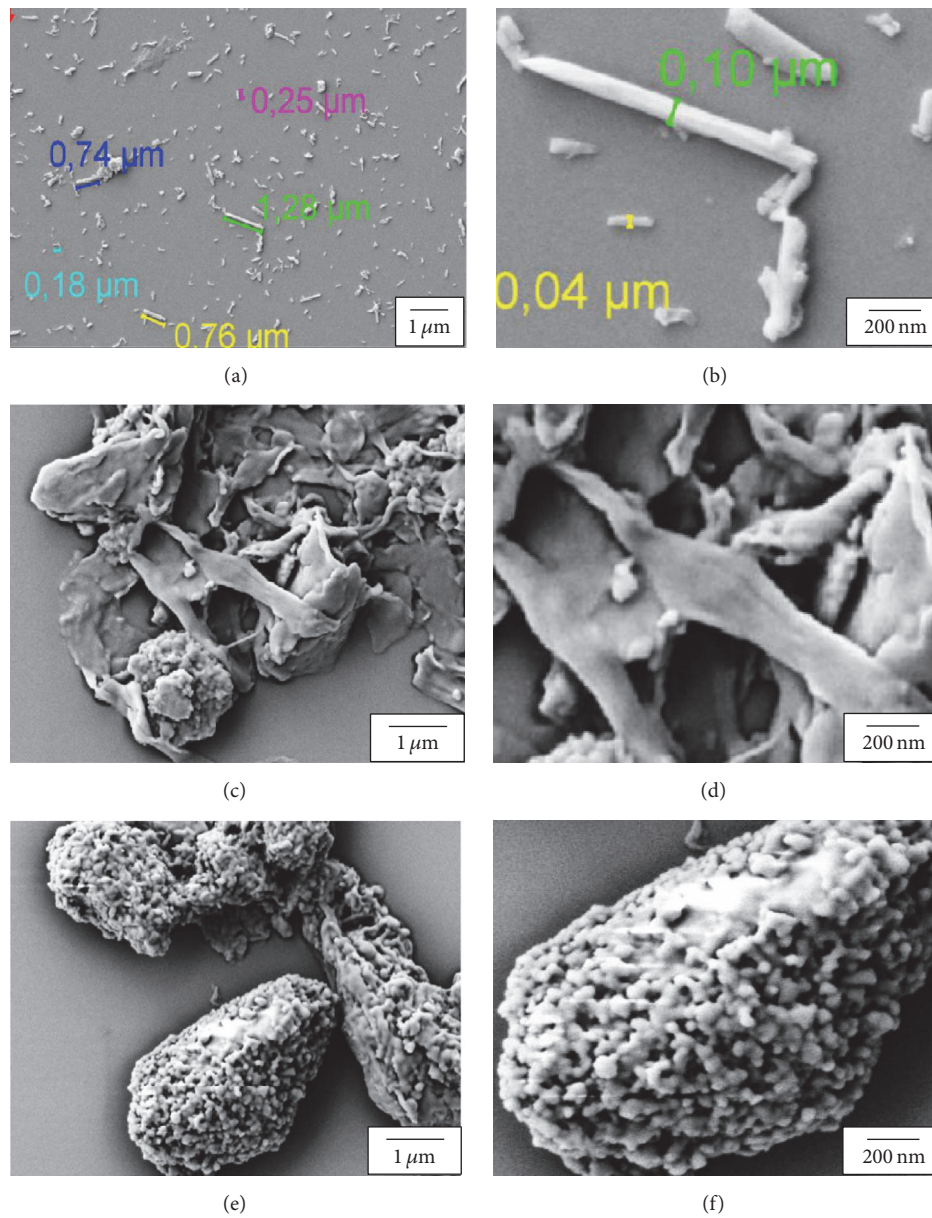


FIGURE 5: SEM micrographs of pristine HNTs (a, b); (c, d) SDS modified HNTs; and (e, f) BCP modified HNTs at two different magnifications.

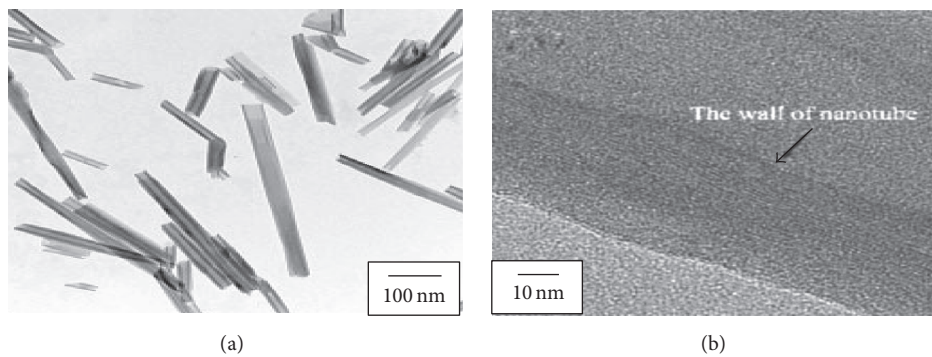


FIGURE 6: TEM micrographs of pristine HNTs at two different magnifications.

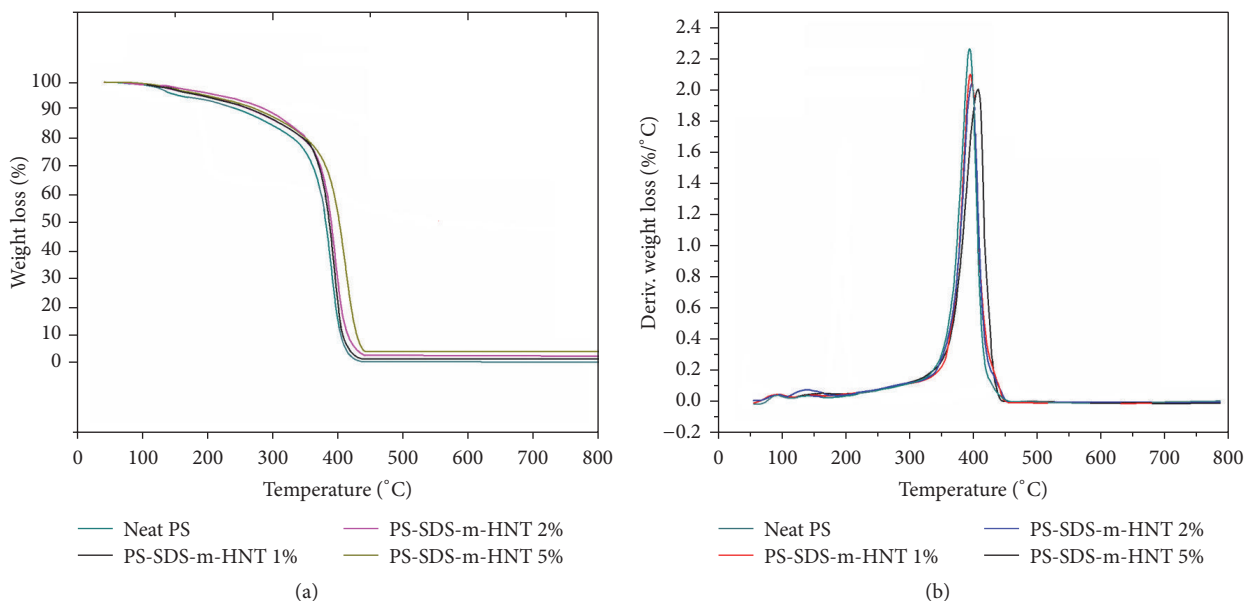


FIGURE 7: (a) TGA and (b) DTG of PS-HNT nanocomposites with SDS modified HNTs.

TABLE 1: Temperatures corresponding to onset of degradation ($T_{0.1}$) and temperatures at 50% degradation ($T_{0.5}$) with respect to wt.% of SDS modified HNTs in the PS-HNT nanocomposite films.

wt.% SDS modified HNT	$T_{0.1}$ (°C)	$T_{0.5}$ (°C)
	250	385
1	262	394
2	276	400
5	289	410

TABLE 2: Temperatures corresponding to onset of degradation ($T_{0.1}$) and temperatures at 50% degradation ($T_{0.5}$) with respect to wt.% of BCP modified HNTs in the PS-HNT nanocomposite films.

wt.% BCP modified HNT	$T_{0.1}$ (°C)	$T_{0.5}$ (°C)
	250	385
1	257	390
2	265	395
5	282	400

films is higher than that of pure PS film and it increases with the increase of the SDS modified HNT loadings in the nanocomposite films. The entrapment of decomposition products by the HNT lumen may be responsible for the improved thermal degradation of PS nanocomposite films.

The results of TGA and DTG analysis of PS nanocomposites containing BCP modified HNTs are shown in Figures 8(a) and 8(b), respectively. It can be clearly seen that the thermal degradation (corresponding to $T_{0.1}$) in the case of pure PS film starts at 250°C, whereas for the nanocomposite films containing 1, 2, and 5% BCP modified HNTs at 257°C, 265°C, and 282°C, respectively (Table 2). Also the temperature corresponding to 50% degradation ($T_{0.5}$) in the case of PS film is found to be 385°C, whereas for nanocomposite films

containing 1, 2, and 5% BCP modified HNTs at 390°C, 395°C, and 400°C, respectively. Hence, it can be concluded that the nanocomposite films show improved thermal stability compared to pure PS film for potential reasons explained above.

Figure 9 shows the TEM micrographs at two different magnifications (10 K and 16 K mag.) for the PS-HNT nanocomposites filled with 1 wt.% of pristine (Figures 9(a) and 9(d)) as well as SDS (Figures 9(b) and 9(e)) and BCP (Figures 9(c) and 9(f)) modified HNTs. The images reveal that the nanocomposite films containing modified HNTs exhibit better dispersion in the PS matrix compared to the pristine HNTs. However, local aggregation of HNTs is also seen at some points, and this aggregation is more pronounced in the case of BCP modified HNTs. Therefore, it can be concluded that SDS modified HNTs give better global distribution of HNTs in the polymer matrix.

Figure 10(a) shows the UV-vis spectra of PS-HNT nanocomposite films with different HNT loadings as a means to study their optical transparency. It can be observed that the onset of absorbance for pure PS film, as well as for the nanocomposite films, is in the range of 200–300 nm. The absorbance values for all the films are very low in the visible region, which indicates that the prepared nanocomposite films do not possess any color and can readily be used as transparent UV-absorbing materials. The absorbance of all the nanocomposite films is slightly higher than that of pure PS film. Also, as the amount of BCP and SDS modified HNTs increases from 1 to 5 wt.%; the absorbance increases in the UV as well as in the visible range. It should be also mentioned that the absorbance values of nanocomposite films containing 1, 2, and 5% BCP modified HNTs are higher than that of nanocomposite films containing 1, 2, and 5% SDS modified HNTs. This is due to presence of π -electrons present in the BCP (PS-b-P4VP). More specifically, molecules containing π -electrons or nonbonding electrons (n-electrons) can

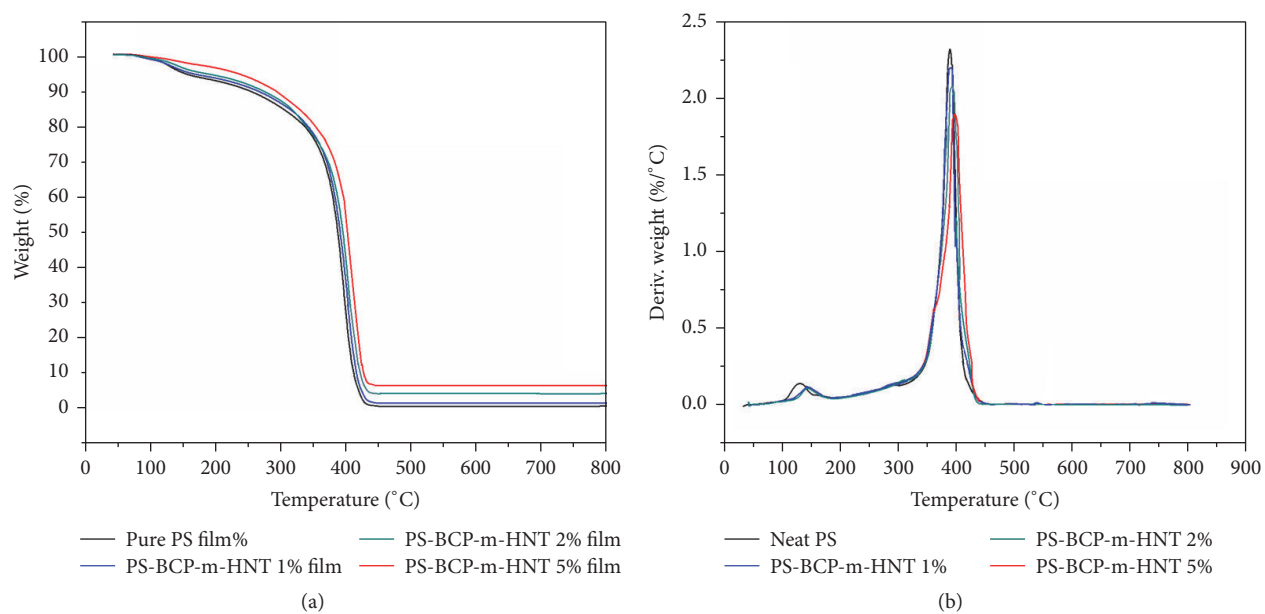


FIGURE 8: (a) TGA and (b) DTG of PS-HNT nanocomposites with BCP modified HNTs.

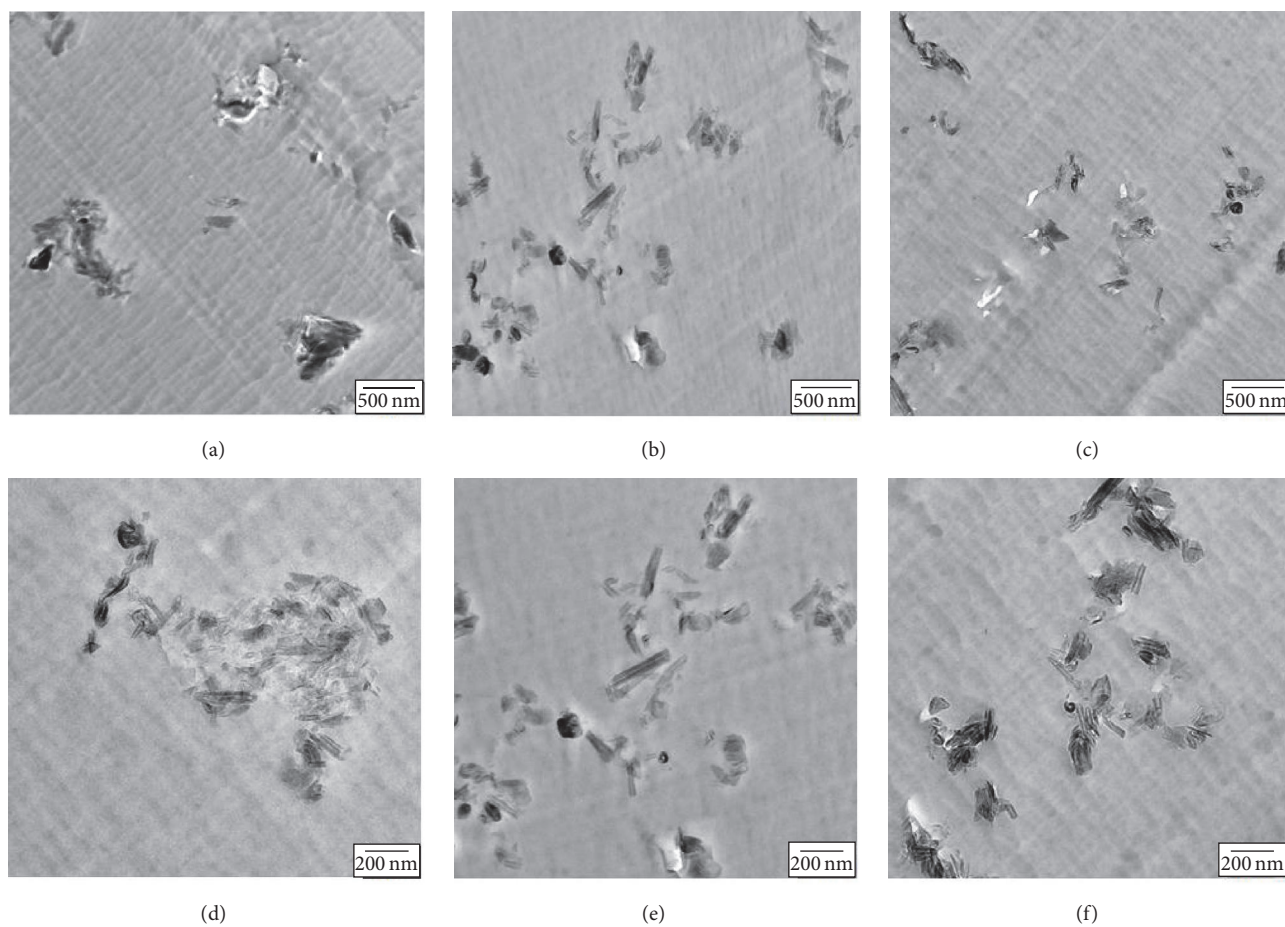


FIGURE 9: TEM micrographs at two different magnifications of PS-HNT nanocomposites containing (a) pristine, (b) SDS, and (c) BCP modified HNTs, respectively, at 1 wt.% loadings.

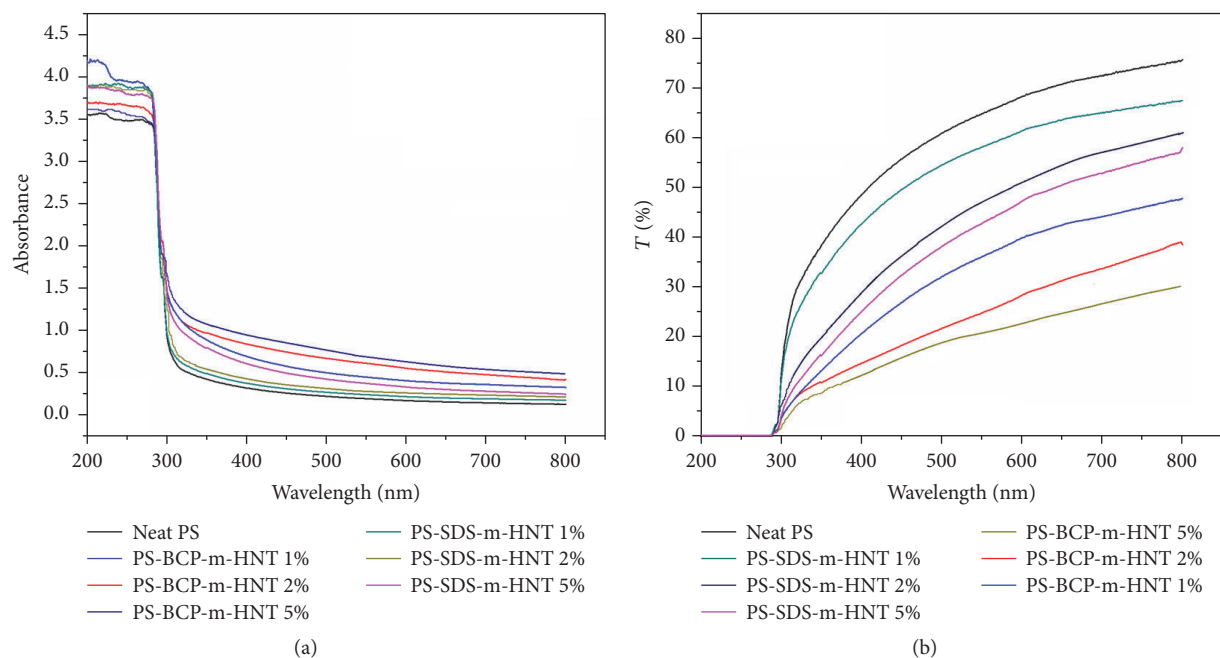


FIGURE 10: (a) UV-vis spectra of PS-HNT nanocomposite films with BCP modified HNTs and SDS modified HNTs compared to pure PS film; (b) % transmittance of PS-HNT nanocomposite films with BCP modified HNTs and SDS modified HNTs compared to neat PS film.

TABLE 3: Variation of % T in case of pure PS film and the nanocomposite films containing SDS modified HNTs with the amount of HNTs (no variations in the optical properties for SDS/PS films have been found compared to neat PS).

% HNT	% T in UV region (350 nm)	% T in visible region (550 nm)
0	38.28	64.72
1	32.66	58.63
2	19.56	45.10
5	16.17	42.73

absorb the energy in the form of ultraviolet or visible light to excite these electrons to higher antibonding molecular orbitals.

Figure 10(b) shows the variation of % T (optical transmittance) of the nanocomposite films as a function of wavelength. It is observed that the transmission starts from 300 nm onwards and it decreases for all nanocomposites films compared to pure PS film. The variation of % T with the amount of HNTs in the nanocomposite films is also studied.

Tables 3 and 4 give the exact data of variation of % T with wt.% loading of modified HNT in the nanocomposite films at two different wavelengths: 350 nm (corresponding to UV region) and 550 nm (corresponding to visible range). It is observed that the % transmittance decreases as % BCP modified HNTs and SDS modified HNTs in the nanocomposites films increase.

By comparing the values in Tables 3 and 4, it can be concluded that the decrease in % T of nanocomposite films with respect to pure PS films is more in the case of films containing

TABLE 4: Variation of % T in case of pure PS film and the nanocomposite films containing BCP modified HNTs with the amount of HNTs (no variations in the optical properties for BCP/PS films have been found compared to neat PS).

% HNT	% T in UV region (350 nm)	% T in visible region (550 nm)
0	38.27	64.71
1	13.18	36.05
2	10.82	24.78
5	8.40	20.01

BCP modified HNTs than that of films containing SDS modified HNTs. Therefore, in order to retain the transparency of nanocomposites films, the SDS modification of HNTs proves to be better than the BCP modified HNTs.

4. Conclusions

In this study, HNTs were successfully modified by SDS and BCP (PS-b-P4VP) organic materials via noncovalent bonds. This facilitated the reduction of inherent hydrophilicity of HNTs making them compatible with the PS polymer matrix. The modifications of HNTs are found to give good dispersion of nanofiller, that is, HNT into the host polymer matrix as shown by the TEM analysis. Hence, the incorporation of modified HNTs in PS matrix for the fabrication of PS nanocomposite films was very successful. A reasonable increase in thermal and optical properties of PS films was observed after loading with small amounts of modified HNTs. The effect of increasing the amount of HNTs in nanocomposite

films on these properties was also studied. It is found that as the loading of HNTs was increased in the PS films, the thermal stability increased without losing much of optical transparency. Since the nanocomposite films are highly transparent and possess good absorption of light in the UV region, these nanocomposites films can be used as transparent and UV shielding materials (greenhouses, food packaging, optical lenses, etc.). Also, comparing the dispersion of HNTs in host matrix and the thermal and optical properties of the nanocomposite films, we conclude that the nanocomposite films containing SDS modified HNTs show better results in all aspects compared to films containing BCP modified HNTs. The modification methods proposed in this study may help in exploiting the potential of HNTs in other polymer matrices due to the achieved HNT dispersion.

Conflicts of Interest

The authors declare that there are no conflicts of interest regarding the publication of this paper.

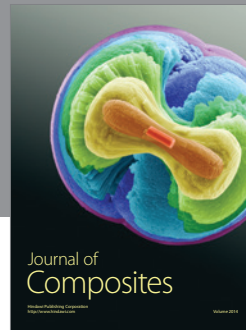
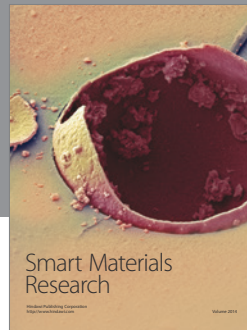
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