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Solubility investigations in the amorphous calcium magnesium carbonate system†

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Amorphous precursors are known to occur in the early stages of carbonate mineral formation in both biotic and abiotic environments. Although the Mg content of amorphous calcium magnesium carbonate (ACMC) is a crucial factor for its temporal stabilization, to date little is known about its control on ACMC solubility. Therefore, amorphous $Ca_xMg_{1-x}CO_3{\cdot}nH_2O$ solids with $0 \le x \le 1$ and $0.4 \le n \le 0.8$ were synthesized and dispersed in $MgCl_2$–$NaHCO_3$ buffered solutions at 24.5 ± 0.5 °C. The chemical evolution of the solution and the precipitate clearly shows an instantaneous exchange of ions between ACMC and aqueous solution. The obtained ion activity product for ACMC (IAP_{ACMC} = "solubility product") increases as a function of its Mg content ($[Mg]_{ACMC}$ = (1 − *x*) × 100 in mol%) according to the expression: $\log(IAP_{ACMC})$ = 0.0174 (±0.0013) × $[Mg]_{ACMC}$ − 6.278 (±0.046) (R^2 = 0.98), where the $\log(IAP_{ACMC})$ shift from Ca (−6.28 ± 0.05) to Mg (−4.54 ± 0.16) ACMC endmember, can be explained by the increasing water content and changes in short-range order, as Ca is substituted by Mg in the ACMC structure. The results of this study shed light on the factors controlling ACMC solubility and its temporal stability in aqueous solutions.

Introduction

Amorphous calcium carbonate (ACC) is a highly soluble solid phase that commonly occurs as a precursor of crystalline $CaCO_3$ (*e.g.* calcite, aragonite) in modern biotic and abiotic precipitation environments. Its (trans)formation has been observed in calcifying organisms,[1–8] microbialites from alkaline lakes,[9,10] biofilms from hot springs[11] and speleothems.[12] ACC can include significant amounts of $MgCO_3$, which is known to play a significant role in its temporal stabilization.[8,10] In this context, the transformation of amorphous calcium magnesium carbonates (ACMC) into crystalline Ca–Mg carbonate minerals is of particular interest as it represents an energetically favorable pathway for high magnesian-calcite (Mg-calcite) and disordered dolomite formation.[13–17] However, the impact of Mg on the temporal stability and transformation behavior of ACMC is not fully elucidated and a systematic study on the solubility of the ACMC system is still missing.

The earlier experimental work of Brečević and Nielsen[18] showed that Mg-free ACC is significantly more soluble than calcite. These authors also documented an ion activity product of ACC ($\log(IAP_{ACC})$; "solubility" of ACC) of −6.39 ± 0.02 at 25 °C. More recently, Gebauer *et al.*[19] found the solubility of ACC to be lower and reported $\log(IAP_{ACC})$ values of −7.51 and −7.42 for ACC phases with a "calcite-like" and a "vaterite-like" short range order, respectively. To our knowledge the only published values so far for ACMC are those obtained in our previous study (Purgstaller *et al.*[15]), where we documented $\log(IAP_{ACMC})$ values ranging from −6.14 to −7.01 for ACMC with ≤10 mol% Mg.

In the scope of an energetic study on synthetic ACMCs, Radha *et al.*[20] observed that ACMC with low Mg content transforms faster into crystalline carbonates than ACMC with high Mg content under atmospheric conditions. The longer stability of ACMC with elevated Mg content has been attributed to the strong bond between structural water and Mg ions which retards the process of dehydration and subsequent phase transformation.[20,21] However, inconsistent data are available for the water contents of ACMC at distinct Mg concentrations.[20–22] For example, the moles of water per unit formula $Ca_xMg_{1-x}CO_3$ are reported to vary from 0.81 to 1.75 for ACMC with 45 ± 2 mol% Mg.[20,21]

New findings indicate that the stability of Mg-free ACC under atmospheric conditions is significantly controlled by the availability of H_2O in the transformation environment. In this regard, Konrad *et al.*[23] proposed that H_2O molecules are adsorbed by ACC promoting its transformation into calcite/vaterite *via* a dissolution and reprecipitation mechanism.

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More recently, experimental studies on ACMC (trans)formation in an aqueous solution suggested that the chemical composition of the reactive solution (*e.g.* Mg/Ca ratio) significantly affects the temporal stability of ACMC and its subsequent transformation into the final product (Mg-calcite, monohydrocalcite *etc.*).[24–27] However, the exact mechanisms controlling ACMC stability and transformation behavior is still under debate and further experimental work is needed to describe the interaction between solutions and ACMC phases.

The aim of the present work is to provide an advanced understanding about the "solubility product" of ACMC as a function of its Mg content in terms of an ion activity product at chemical steady-state conditions. Therefore, distinct ACMCs ($Ca_xMg_{1-x}CO_3{\cdot}nH_2O$; $0 \leq x \leq 1$ and $0.4 \leq n \leq 0.8$) were synthesized and subsequently dispersed in $MgCl_2$–$NaHCO_3$ buffered solutions in order to assess i) the elemental exchange of ACMC with the aqueous phase and ii) the factors controlling ACMC solubility and its temporal stability in aqueous solutions.

Experimental section

Synthesis of amorphous calcium magnesium carbonates

Amorphous calcium magnesium carbonates with Mg contents ranging from 0 to 100 mol% Mg (hereafter referred to as ACMCs) were synthesized by a previously described method.[23] Briefly, 80 mL of a 250 mM $(Ca,Mg)Cl_2$ solution was poured into a beaker containing 80 mL of a 250 mM Na_2CO_3 solution. The reaction products were immediately separated from solution by a 0.2 μm cellulose filter using a suction filtration unit. Subsequently, the separated precipitate was washed with ultrapure water (Millipore Integral 3: 18.2 MΩ cm^{-1}) and transferred into a freeze dryer (Virtis Benchtop 3L). The freeze-dried ACMCs were stored in closed vials in a desiccator with silica gel (relative humidity = 3%). In total, 9 synthesis experiments were carried out, where the Mg content of the $(Ca,Mg)Cl_2$ stock solution, $[Mg]_{stock} = \{[Mg]/([Ca] + [Mg])\} \times 100$, was systematically varied between 0 and 100 mol% (see Table 1). All solutions were prepared by analytical grade chemicals ($CaCl_2{\cdot}2H_2O$, $MgCl_2{\cdot}6H_2O$ and Na_2CO_3 from Roth) mixed with ultrapure water.

Experimental setup for solubility study

The chemical composition of the synthesized ACMC solids used for solubility investigations is shown in Table 1. Experiments were performed at 24.5 ± 0.5 °C in a 100 mL glass reactor containing 50 mL of a 100 mM $NaHCO_3$ and 30 mM $MgCl_2$ solution, stirred at 350 rpm. The pH of the solution was adjusted to 8.33 ± 0.03 by the addition of a 500 mM NaOH solution. At the onset of the experiment, 1.5 g of the synthetic ACMC sample was introduced into the reactor. The temporal evolution of the mineralogy was monitored using an *in situ* Raman probe immersed in the suspension. In order to follow the chemical composition of the solution and the solid phase, homogeneous sub-samples of the experimental solution/suspension (1.5 mL) were collected with a pipette at certain reaction times (Table S1†). The solids were separated from the solution by a 0.2 μm cellulose acetate filter using a suction filtration unit, washed with ethanol and dried in an oven at 40 °C.

Solid phase characterization

X-ray diffraction (XRD) patterns of the synthesized ACMCs (Fig. S1†) and of the ACMCs separated from the experimental solutions during the solubility study (not shown) were acquired using a PANalytical X'Pert Pro diffractometer (Co-Kα radiation) at a 2Θ range from 4 to 85° and a scan speed of 0.03° s^{-1}. Thermogravimetric analyses (TGA) of freeze-dried ACMCs (Fig. S2†) were realized using a PerkinElmer STA8000. The samples were heated from 25 °C to 800 °C at 10 °C min^{-1} in the presence of 99.999% N_2 atmosphere. Selected synthesized ACMCs were gold coated and imaged (Fig. S3†) using a scanning electron microscope (SEM, ZEISS DSM 982 Gemini). Time-resolved *in situ* Raman spectroscopy of the experimental solution/suspension was realized using a Raman RXN2™ analyzer from Kaiser Optical Systems with a Kaiser MR Probe head (quarter-inch immersion optic) and a 785 nm laser beam. *In situ* Raman spectra were collected every 60 s in the 100–1890 cm^{-1} region with a resolution of 1 cm^{-1}.

Table 1 Chemical composition of synthesized ACMCs

Sample	$[Mg]_{stock}$[a] (mol%)	$[Mg]_{ACMC}$[b] (mol%)	$nH_2O_{p.f.}$[c] (mol)	Sample composition
ACC	0	0	0.44	$CaCO_3{\cdot}0.44H_2O$
ACMC_9	20.1	9.0	0.49	$Ca_{0.91}Mg_{0.09}CO_3{\cdot}0.49H_2O$
ACMC_15	30.3	14.9	0.51	$Ca_{0.85}Mg_{0.15}CO_3{\cdot}0.51H_2O$
ACMC_22	40.1	21.9	0.53	$Ca_{0.78}Mg_{0.22}CO_3{\cdot}0.53H_2O$
ACMC_31	50.1	30.8	0.56	$Ca_{0.69}Mg_{0.31}CO_3{\cdot}0.56H_2O$
ACMC_39	60.7	39.4	0.59	$Ca_{0.61}Mg_{0.39}CO_3{\cdot}0.59H_2O$
ACMC_53	72.1	53.4	0.66	$Ca_{0.47}Mg_{0.53}CO_3{\cdot}0.66H_2O$
ACMC_80	89.1	80.0	0.71	$Ca_{0.20}Mg_{0.80}CO_3{\cdot}0.71H_2O$
AMC	100	100	0.79	$MgCO_3{\cdot}0.79H_2O$

[a] Mg content of the stock solution in mol%. [b] Mg content of the ACMC solid in mol%. [c] Moles of water per formula unit $Ca_xMg_{1-x}CO_3$, where $(1 - x) = [Mg]_{ACMC}/100$.

Chemical composition of the experimental solutions and solids

During the experimental run, the pH of the experimental solution was measured with a SI Analytics Silamid® gel electrode, which was calibrated against NIST buffer standard solutions at pH 4.01 and 7.00. The total alkalinity of the experimental solutions was measured by a Schott TitroLine alpha plus titrator using a 10 mM HCl solution with a precision of ±2%. The aqueous Ca, Mg and Na concentrations of the $(Ca,Mg)Cl_2$ stock solutions, of the experimental solutions and of the solids (digested in 6% HNO_3) were determined using inductively coupled plasma optical emission spectrometry (Perkin Elmer Optima 8300 DV). The analytical error was <±3% for Ca and Mg analyses and <±5% for Na analyses.

Aqueous speciation and ion activity product calculation

The aqueous speciation of the experimental solutions was calculated at 25 °C using the PHREEQC software together with its minteq.v4 database. For the ionic strength of our experiments (*i.e.* 0.17 ± 0.02 M; Table S2†), the calculation of individual ion activity coefficients for the solute species is based on the Davis equation.[28,29] The activities (a) of Ca^{2+}, Mg^{2+} and CO_3^{2-} ions in solution and the stoichiometry of the digested $Ca_xMg_{1-x}CO_3{\cdot}nH_2O$ solids (where $(1-x) = [Mg]_{ACMC}/100$) were used to calculate ion activity products for the amorphous calcium magnesium carbonates (IAP_{ACMC}) as a function of experimental time according to equation

$$IAP_{ACMC} = (aCa^{2+})^x(aMg^{2+})^{1-x}(aCO_3^{2-})(aH_2O)^n \quad (1)$$

In the present solubility experiments, the aH_2O values of the experimental solutions remained constant at 0.955 ± 0.004. Thus, for the calculation of the IAP_{ACMC}, the aH_2O was assumed to be unity.

Results and discussion

Characterization of synthesized ACMC material

The measured Mg contents of the stock solutions, $[Mg]_{stock}$, and of the synthesized ACMC samples, $[Mg]_{ACMC}$, are reported in Table 1 and displayed in Fig. 1. The results reveal a preferential enrichment of the amorphous solid in Ca and are in good agreement with those of Radha *et al.*,[20] who synthesized a set of amorphous $Ca_xMg_{1-x}CO_3{\cdot}nH_2O$ solids by batch method using $(Ca,Mg)Cl_2$ and Na_2CO_3 solutions (Fig. 1). The preferential enrichment of the amorphous phase in Ca likely stems from the strong free energy of hydration of the Mg^{2+} ion compared to Ca^{2+}.[20,30] The strongly hydrated aqueous Mg^{2+} is a well-known limitation in the formation of anhydrous crystalline Ca–Mg-carbonates, such as Mg-calcite, dolomite and magnesite, at ambient temperatures.[31] Instead, hydrous crystalline Mg-carbonates (*e.g.* hydromagnesite)[32] and hydrous ACMC phases[33,34] tend to precipitate.

Representative thermogravimetric analysis (TGA) and differential scanning calorimetry (DSC) curves of the synthesized ACMCs are presented in Fig. S2 (ESI†). The DSC curve of ACC (Ca endmember of ACMC) shows an endothermic peak at 120 °C, a sharp exothermic peak at 340 °C and a second endothermic peak at 731 °C (Fig. S2A†). These DSC features have been earlier shown to be associated with the enthalpies of dehydration, crystallization and decomposition of ACC.[20,35] In contrast to ACC, AMC (Mg endmember of ACMC) decomposes without crystallization, as it is indicated by the absence of the sharp exothermic peak in the DSC curve (Fig. S2F†). The weight loss curves of ACC and AMC show two weight loss steps (Fig. S2A and F†), where the first step is due to the loss of structural water and the second step due to the decomposition of $CaCO_3$ and $MgCO_3$ to CaO, MgO and CO_2.[20,35] After ACC dehydration, the TGA curve achieves a plateau in heating before $CaCO_3$ decomposes to CaO and CO_2 at near 600 °C (Fig. S2A†). In contrast, AMC shows a continuous drop of the TGA curve (without a plateau) for the same temperature range, which indicates that AMC dehydration and decomposition may overlap (Fig. S2F†). Based on our TGA data we suggest that AMC decomposition occurs at around 275 °C, as it is indicated by the inflection point of the weight loss curve (Fig. S2F†). In contrast to the pure endmembers, the ACMCs show multistep weight loss curves (Fig. S2B–E†). These trends are in accordance with the findings of Radha *et al.*,[20] who associated these features with multistep carbonate decomposition caused by (i) an initial heterogeneous amorphous material or (ii) thermally induced phase segregation during the heating process.

The water contents of the synthetic ACMCs were calculated from the weight losses between 25 and 275 °C. Above this temperature, mass loss might occur due to the decomposition of $MgCO_3$ to MgO, as it is suggested by the TGA curve

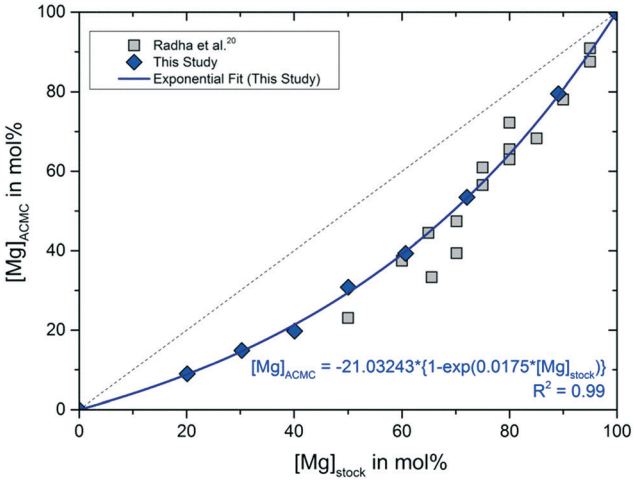

Fig. 1 Mg content of the stock colutions, $[Mg]_{stock}$, *versus* Mg content of the ACMC precipitates, $[Mg]_{ACMC}$, and respective values reported by Radha *et al.*[20] The increase of $[Mg]_{ACMC}$ as a function of $[Mg]_{stock}$ of this study can be described by the equation $[Mg]_{ACMC} = -21.03243 \pm 2.1968 \times \{1 - \exp(0.0175 \pm 0.0009 \times [Mg]_{stock})\}$ (R^2 = 0.99; solid line). Analytical uncertainties are included in the symbol size.

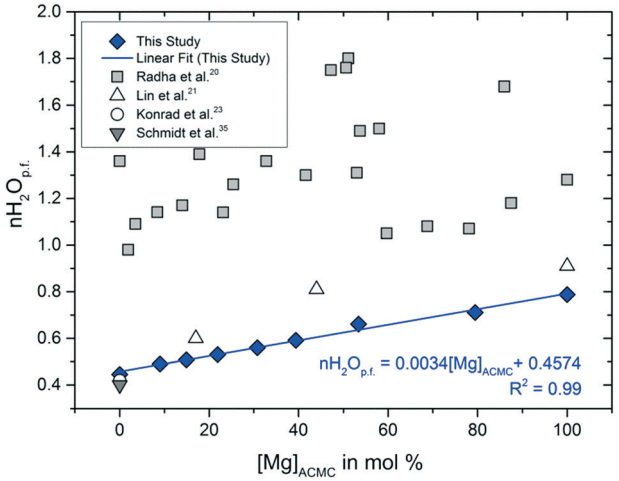

Fig. 2 Mg content of ACMC, $[Mg]_{ACMC}$, *versus* the moles of water per unit formula $Ca_xMg_{1-x}CO_3$, $nH_2O_{p.f.}$, from synthesis experiments conducted in the present study and from literature. Analytical uncertainties are included in the symbol size.

of AMC (Fig. S2F†). The moles of water per unit formula $Ca_xMg_{1-x}CO_3$ ($nH_2O_{p.f.}$) determined by thermal analyses vary between 0.4 for ACC and 0.8 for AMC (Table 1, Fig. 2). As it can be seen in Fig. 2, the water content determined for ACC is in excellent agreement with those reported by Schmidt *et al.*[35] and Konrad *et al.*[23] Overall, thermal analyses revealed a linear correlation between the $nH_2O_{p.f.}$ and the $[Mg]_{ACMC}$ according to the equation

$$nH_2O_{p.f.} = 0.0034 \pm 0.0002[Mg]_{ACMC} + 0.4574 \pm 0.0124 \quad (2)$$

where $R^2 = 0.99$. Previous experimental work documented significantly higher $nH_2O_{p.f.}$ values for ACMCs (Fig. 2).[20] These differences probably originate from different drying methods (freeze-dryer *versus* vacuum oven). In the study of Radha *et al.*,[20] the loosely bound physisorbed H_2O on the ACMC particles were probably not completely removed by the vacuum oven prior to thermogravimetric analyses. However, our findings are in good agreement with the results of Lin *et al.*,[21] who reported $nH_2O_{p.f.}$ values of 0.6, 0.8 and 0.9 for ACMC with 17, 44 and 100 mol% Mg (Fig. 2). In the latter case, the synthesized ACMC samples were washed with ethanol and lyophilized prior to TGA-analysis.

Earlier models on the structure of Mg-free ACC suggested the presence of partially mobile water in the amorphous phase.[36,37] In this context, ACC has been described to consist of a porous Ca-framework with interconnected channels formed by water molecules and carbonate ions, allowing water mobility to a certain extent.[38] More recently, Jensen *et al.*[39] has shown that water molecules are mainly coordinated to Ca and carbonate ions and less frequently to other water molecules, ruling out the scenario of hydrogen-bonded networks. This is consistent with the NMR-results of Lin *et al.*[21] who concluded that the coordination shell of Mg in ACMC contains at least one water molecule. Indeed, in the present study, the $nH_2O_{p.f.}$ increases as a function of the Mg content (Fig. 2), indicating that partially hydrated Mg is incorporated into the amorphous solid. Our obtained results however suggest that the $nH_2O_{p.f.}$ values for ACMC are significantly lower than previously reported.[20,22]

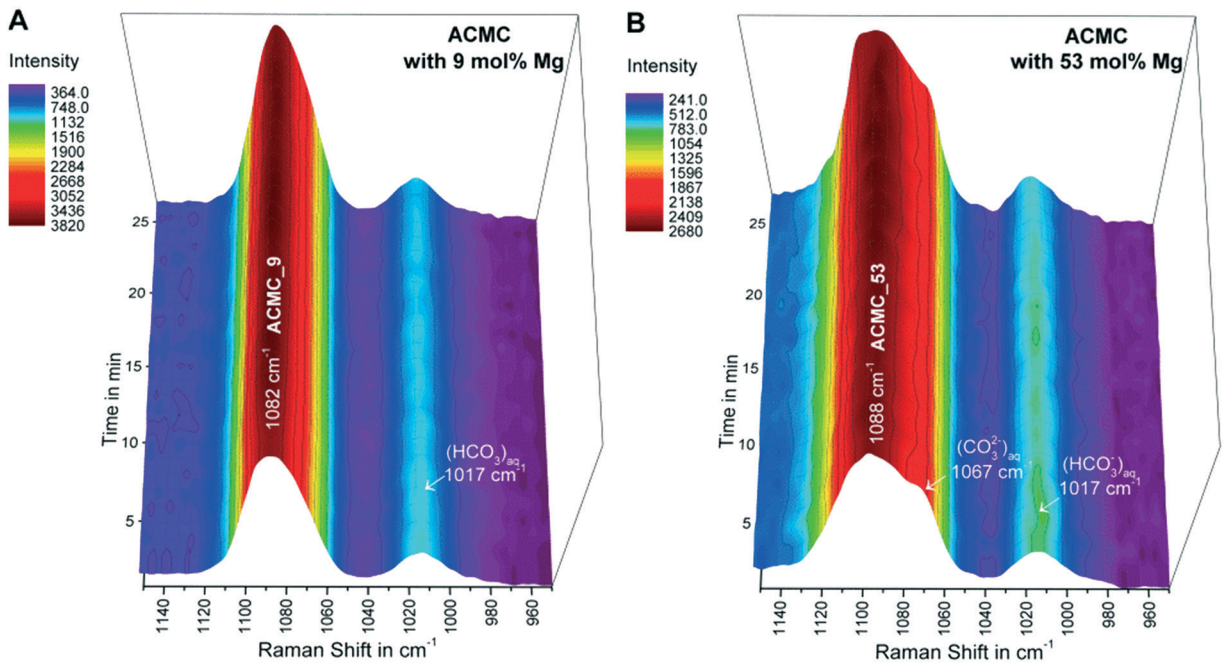

Fig. 3 3D plots of *in situ* Raman spectra showing the temporal evolution of the v_1 band of ACMC for the experiments conducted with ACMC_9 (A) and ACMC_53 (B). The Raman bands at 1067 cm^{-1} and 1017 cm^{-1} correspond to aqueous CO_3^{2-} and HCO_3^-, respectively.

Table 2 Summary of *in situ* Raman results and calculated ion activity product values of solubility experiments

Sample	Raman ν_1 band[a] (cm^{-1})	$Time_{ACMC}$[b] (min)	$[Mg]_{ACMC\_av}$[c] (mol%)	$\log(a{CO_3}^{2-})_{av}$[d]	$\log(a{Ca}^{2+})_{av}$[d]	$\log(a{Mg}^{2+})_{av}$[d]	$\log(IAP_{ACMC})_{av}$[e]
ACC	1080 ± 1	15	2.2 ± 0.2	−3.14 ± 0.01	−3.08 ±0.03	−2.25 ± 0.02	−6.20 ± 0.02
ACMC_9	1082 ± 1	24	10.5 ± 0.2	−3.01 ± 0.02	−3.22 ±0.03	−2.23 ± 0.01	−6.13 ± 0.02
ACMC_15	1083 ± 1	54	15.8 ± 0.3	−2.91 ± 0.01	−3.31 ±0.04	−2.21 ± 0.01	−6.04 ± 0.03
ACMC_22	1084 ± 1	85	21.2 ± 0.7	−2.87 ± 0.03	−3.38 ±0.05	−2.20 ± 0.03	−6.00 ± 0.03
ACMC_31	1085 ± 1	64	29.3 ± 1.6	−2.67 ± 0.02	−3.54 ±0.08	−2.18 ± 0.05	−5.81 ± 0.07
ACMC_39	1086 ± 1	51	37.4 ± 0.5	−2.58 ± 0.02	−3.68 ±0.06	−2.21 ± 0.06	−5.71 ± 0.03
ACMC_53	1088 ± 1	20	51.3 ± 0.6	−2.51 ± 0.02	−3.78 ±0.05	−2.12 ± 0.01	−5.43 ± 0.02
ACMC_80	1095 ± 1	9	78.5 ± 0.4	−2.39 ± 0.03	−4.10 ±0.10	−2.14 ± 0.01	−4.96 ± 0.01
AMC	1098 ± 1	5	100	−2.30 ± 0.01	—	−2.12 ± 0.01	−4.42 ± 0.01

[a] Raman ν_1 peak positions of ACMC solids dispersed in the $MgCl_2$–$NaHCO_3$ buffered solutions. [b] Time interval in which ACMC was stable in the experimental solution. [c] Average Mg content of the ACMC solid at chemical steady state conditions (used for IAP_{ACMC} calculation). [d] Average activities of ${CO_3}^{2-}$, Ca^{2+} and Mg^{2+} ions in the experimental solution at chemical steady state conditions (used for IAP_{ACMC} calculation). [e] Average ion activity product for ACMC calculated according to eqn (1). Note that the input data for speciation calculations and selected output data are listed in Table S2.

Exchange of ions between ACMC and aqueous solution

After the synthesized ACMCs were dispersed in the $MgCl_2$–$NaHCO_3$ buffered solutions, the collected *in situ* Raman spectra of the suspensions revealed the presence of a broad ${CO_3}^{2-}$ symmetric stretch (ν_1 band) of ACMC at 1080–1098 cm^{-1} (*e.g.* experiments ACMC_9 and ACMC_53 in Fig. 3). The ν_1 band shows systematic shifts in peak position (Table 2) and peak broadening due to different Mg contents of the ACMCs. These features arise from changes in the metal–oxygen bond length due to the presence of shorter Mg–O bonds compared to those of Ca–O in the amorphous solid.[15,40] The intensities of the ν_1 bands of the ACMC phases remained constant for experimental times ranging from 5 to 85 min (referred to as $Time_{ACMC}$ in Table 2). After this time interval, the ν_1 bands of the ACMCs decreased in intensity, while the ν_1 bands of crystalline hydrous and/or anhydrous Ca–Mg carbonates (Mg-calcite, monohydrocalcite and/or nesquehonite) evolved (not shown here). Thus, $Time_{ACMC}$ denotes the time interval in which the ACMC was stable in the experimental solution (Table 2). Note that, the *in situ* Raman observations were confirmed by *ex situ* XRD analyses of the collected solid samples.

Although, $Time_{ACMC}$ increased as a function of the Mg content in experiments conducted with ACMCs containing ≤22 mol% Mg, a reverse trend was obtained in experiments with ACMCs containing >22 mol% (Table 2). These observations imply that the temporal stability of ACMC in a solution is not strictly controlled by its primary Mg content. In contrast, the obtained data revealed changes in the chemical composition of the reacting ACMC and the experimental solution which might play a crucial role in the temporal stabilization of the precursor phase and its subsequent transformation into distinct crystalline calcium magnesium carbonates.

The chemical composition of the experimental solutions and solids collected during $Time_{ACMC}$ can be found in Table S1.† Exemplarily, the temporal evolution of the chemical composition of the solutions and solids in experiments conducted with ACMC_53 (53 mol% Mg) and with ACMC_9 (9 mol% Mg) is illustrated in Fig. 4A–C. The obtained results documented a significant change in the chemical composition of the solids and experimental solutions within about 10 s after the ACMCs were dispersed into the $MgCl_2$–$NaHCO_3$ solutions. In each experiment, chemical steady state conditions were attained at a reaction time of about 2 min (Fig. 4A–C; Table S1†). The fast reaction between the amorphous solid and the solution can be explained by the nano-porous structured ACMC material consisting of large and highly reactive surface areas (Fig. S3†). The shifts in pH and alkalinity of the $MgCl_2$–$NaHCO_3$ buffered solution increase with increasing Mg content of the synthesized ACMCs (Fig. 4D). For example, dispersion of ACMC with 9 mol% Mg (ACMC_9) yields a slight shift in pH and alkalinity from 8.3 to 8.5 and 104 to 107 mM, respectively, whereas for ACMC with 53 mol% Mg (ACMC_53) a significant shift in pH to 9.1 and alkalinity to 143 mM was observed (Fig. 4A and B). This indicates that for attainment of chemical equilibrium between solid and solution, larger amounts of ACMC are required to dissolve in the experiment with ACMC_53 (53 mol% Mg) compared to the experiment conducted with ACMC_9 (9 mol% Mg).

The results revealed two distinctive trends of Mg and Ca exchange between ACMCs and the solution, depending on the Mg concentrations of synthesized ACMCs: (i) in experiments conducted with ACMCs containing ≤15 mol% Mg a net release of Ca into the solution together with an uptake of Mg from the solution into the solid is observed, which results in ACMCs with slightly higher Mg contents compared to the synthesized ACMC (*e.g.* ACMC_9 in Fig. 4B and C; Table S1†). (ii) In contrast, in experiments conducted with ACMCs containing ≥20 mol% Mg a net release of Mg into the experimental solution and lower $[Mg]_{ACMC}$ values compared to the synthesized ACMCs were observed (*e.g.* ACMC_53 in Fig. 4B and C; Table S1†). These observations suggest a dynamic exchange of Me^{2+} (Ca^{2+} and Mg^{2+}) ions between the amorphous solid and the experimental solution. It is likely that the total availability of Me^{2+} (Ca^{2+} and Mg^{2+}) and ${CO_3}^{2-}$ ions in the prevailing solution affects the formation of distinct crystalline calcium magnesium carbonates and the temporal stability of the ACMC precursor in solution. Evidence for that also comes from a recent study on ACC transformation,

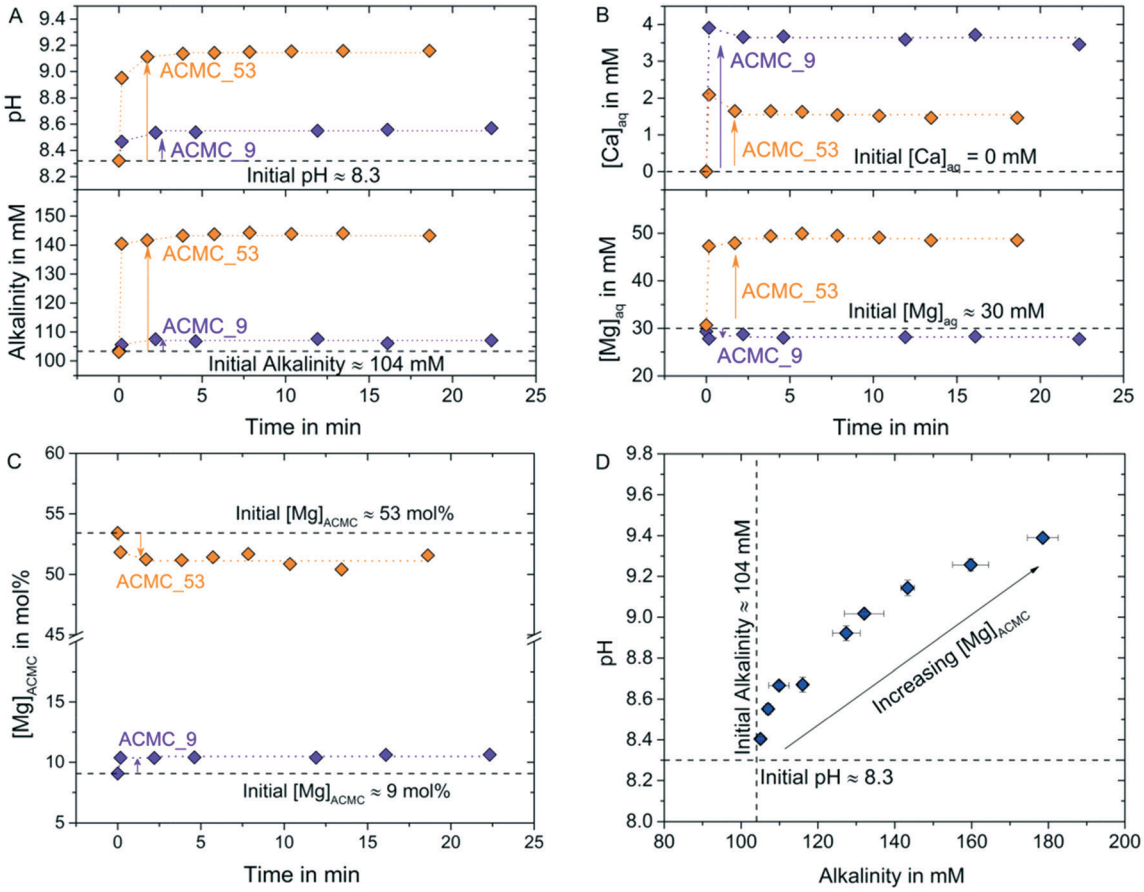

Fig. 4 Temporal evolution of (A) pH and alkalinity concentration, (B) aqueous Ca and Mg concentration of the experimental solution, $[Ca]_{aq}$ and $[Mg]_{aq}$, and (C) Mg content of ACMC solid, $[Mg]_{ACMC}$, after synthesized ACMC solids were dispersed into the $MgCl_2$–$NaHCO_3$ solutions. Exemplarily the data for ACMC_53 and ACMC_9 are shown, where the experimental solution achieves constant chemical composition after 2 min of reaction time. Note that chemical data obtained at >2 min were used for estimating the ion activity products of the distinct ACMCs according to eqn (1) (Table 2). (D) Average alkalinity concentration *versus* pH values of the experimental solution at chemical steady state conditions (>2 min, see Table S2†). The shifts in pH and alkalinity increase with increasing $[Mg]_{ACMC}$.

showing prolonged stability of Mg-free ACC of up to ~11 hours in high concentrated $MgCl_2$ solutions (up to 1000 mM).[26] It was concluded that strong aquocomplex formation of the dissolved carbonate molecule with Mg^{2+} ions retarded mineral formation and increased the stability of ACC in solution.[26]

Solubility approach

The experimental solution in all experimental runs achieved chemical steady state conditions after 2 min of reaction time. These conditions remained steady during the time interval in which solid analyses (Raman spectra and XRD pattern) documented the stability of ACMC in the aqueous phase ($Time_{ACMC}$ in Table 2). Thus, speciation calculations were performed at reaction times ≥2 min and ≤$Time_{ACMC}$. The input data for speciation calculations using PHREEQC modeling and selected output data are summarized in Table S2.† The results showed that the experimental solutions exhibit internal partial pressures of CO_2 ranging from $10^{-1.8}$ to $10^{-2.9}$ atmospheres (see $p_{CO_2(g)}$ in Table S2†). As such, the p_{CO_2} pressure of the experimental solutions is higher compared to the atmosphere and thus CO_2 is likely degassing from the experimental solutions. However, in context of the present study, samples of the suspensions were collected during a short time interval (<45 min) in which $p_{CO_2(g)}$, pH and alkalinity values were monitored to be constant within the analytical precision (Table S2†). Thus, the obtained activities of Ca^{2+}, Mg^{2+} and CO_3^{2-} ions of the solution and the ACMC stoichiometry at the given reaction time were used to calculate individual IAP_{ACMC} (solubility) values according to eqn (1) (Table S2†). The average $\log(IAP_{ACMC})$ values are listed in Table 2 and plotted in Fig. 5 as a function of the Mg content of the ACMC. Note here that the average Mg content of the ACMC solid at chemical steady state conditions (after Ca and Mg ion exchange) was used for IAP_{ACMC} calculations (Table 2) and was plotted in Fig. 5. The results clearly show that the solubility of ACMC increases as a function of the Mg content (Fig. 5). The linear fit of the $\log(IAP_{ACMC})$ values obtained in the context of the present study is in good agreement with the log(IAP) value for ACC documented by Brečević and Nielsen,[18] while the solubility products for ACCs reported by Gebauer *et al.*[19] are significantly smaller. The obtained

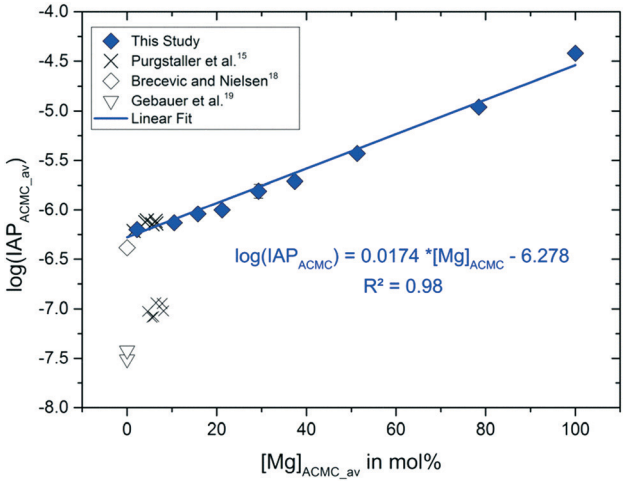

Fig. 5 Average Mg content of ACMC, $[Mg]_{ACMC\_av}$, *versus* the average ion activity product of ACMC, $log(IAP_{ACMC})_{av}$ (resembling the solubility product of the ACMC phase), calculated according to eqn (1) at chemical steady state conditions (see values in Table 2). The $log(IAP_{ACMC})$ values for Mg-free ACC are taken from Brečević and Nielsen[18] and Gebauer *et al.*,[19] while those for ACMC with ≤10 mol% Mg are taken from Purgstaller *et al.*[15] The analytical uncertainties are included in the symbol size.

$log(IAP_{ACMC})$ values of −6.19 ± 0.02 and −6.13 ± 0.02 for ACMCs with 2 and 10 mol% Mg, respectively, lay within the range of −6.14 ± 0.04 for the low-Mg ACMC reported by Purgstaller *et al.*[15] In this respect, it has to be noted that the solids collected in the present experiments were filtered through 0.2 μm membranes. Although the separation yields aggregated spherical nano-particles (see Fig. S4†) and the solutions were visibly clear after filtration, the presence of ACMC individuals in the filtered solutions cannot be completely ruled out at this stage. Partially remaining ACMC in the solution would slightly increase the measured ion concentrations. However, the linear fit of the herein obtained $log(IAP_{ACMC})$ values is in good agreement with the $log(IAP_{ACC})$ value from Brečević and Nielsen,[18] who determined the solubility of ACC by titration of water into the suspension without physical separation of the solid phase, indicating the effect of ACMC individuals in the filtered solutions to be negligibly small. The positive correlation between the $log(IAP_{ACMC})$ and $[Mg]_{ACMC}$ values obtained from this study, from Brečević and Nielsen[18] and Purgstaller *et al.*[15] can be described by the expression:

$$log(IAP_{ACMC}) = 0.0174 \pm 0.0013 \times [Mg]_{ACMC} - 6.278 \pm 0.046 \quad (3)$$

where $R^2 = 0.98$.

Note here that the calculation of the IAP_{ACMC} values is based on the assumption that the ACMCs are amorphous Ca–Mg carbonate solid-solutions. However, a previous study on synthetic ACMCs by Radha *et al.*[20] points towards a phase separation of ACMC solids with ≥47 mol% Mg. It has been suggested that ACMC with ≥47 mol% are composed by a mixture of a material nearly pure AMC with a material nearly 50 mol% Mg. However, such heterogeneity may also be induced during the heating of the ACMC solids during thermogravimetric analyses (thermally driven $MgCO_3$ segregation).[20] More recently, the ^{13}C-NMR measurements on ACMCs by Yang *et al.*[41] showed two types of carbonate ions in synthetic ACMC solids, whose short-range orders are identical to those of ACC and AMC. Based on these observations, Yang *et al.*[41] suggested that the ACMC solids comprise a homogeneous mixture of the nano-clusters of ACC and AMC. In order to test the validity of the Ca–Mg carbonate solid-solution system, we calculated the IAP values of the experimental solutions at chemical steady state conditions with respect to the endmembers ACC and AMC according to the equations $IAP_{ACC} = (aCa^{2+})\cdot(aCO_3^{2-})$ and $IAP_{AMC} = (aMg^{2+})\cdot(aCO_3^{2-})$, respectively. As the Mg content of the solid increases from 2 to 78 mol% Mg, the $log(IAP_{AMC})$ and $log(IAP_{ACC})$ values change from −5.4 to −4.5 and −6.2 to −6.5, respectively. If the ACMCs consist of an ideal two phase mixture of ACC and AMC, we would expect constant $log(IAP_{ACC})$ and $log(IAP_{AMC})$ in all experiments, irrespective of the Mg content of the solid initially introduced into the reactor. As such, the presence of discrete AMC and ACC phases in the ACMC solid can be excluded. In contrast, the obtained linear correlation between the average $log(IAP_{ACMC})$ and $[Mg]_{ACMC}$ values (see Fig. 5) points toward the presence of an homogeneous single phase in each experimental run.

It is also worth pointing out that the ^{13}C-NMR measurements on ACMCs by Yang *et al.*[41] indicated the existence of bicarbonate species embedded in the matrix formed by ACC and AMC. Moreover, based on the Mg isotope composition of Mg-ACCs, Mavromatis *et al.*[42] suggested that during the very fast precipitation of the amorphous phase from a highly supersaturated solution, $MgHCO_3^+$ and $MgCO_3^0$ species might be directly incorporated into the precipitating solid together with $Mg(H_2O)_6^{2+}$. If this holds true, the activities of Ca and Mg carbonate species have to be considered in the calculation of the IAP_{ACMC} values. For this purpose, however, a more detailed characterization of the structure of ACMC is required, which should be in the focus of future studies.

Although the solubility of ACC is about two orders of magnitude higher than that of calcite, the dependence of IAP_{ACMC} on the Mg content of the solid (Fig. 5) is somehow similar to that observed for the Mg-calcite system.[43] Indeed experimental work conducted during the last 3 decades documents an overall increase in solubility of Mg-bearing calcites with increasing Mg content.[43–45] Note, that significant discrepancies exist between reported solubility values of Mg-calcites, which were attributed to differences in mineral source and structure (biogenic or synthetic solids) as well as different experimental setups.[43–45] The change in calcite solubility as Mg is incorporated in its structure can likely be attributed to the distortion of the octahedral site in calcite. The substitution of Ca by Mg in the calcite structure results in a decrease of interatomic metal–O distances and shortening of the *c*-axis.[46,47] Consequently, Ca, C and O atoms exhibit large thermal motions which cause periodic stretching and weakening of the bonds in the structure.[46]

In contrast to the crystalline Ca–Mg carbonate system, amorphous carbonates exhibit no order beyond 1.5 nm.[40,48] *In situ* Raman results from the present work and from previous studies[15,40] however showed that the substitution of Ca^{2+} by Mg^{2+} changes the average metal–oxygen bond lengths in the amorphous phase, which might change its short-range order and solubility. Moreover, the linear relationship between the IAP_{ACMC} and $[Mg]_{ACMC}$ values (Fig. 1) as well as of the $[Mg]_{ACMC}$ and $nH_2O_{p.f.}$ values (Fig. 2) suggests that the increase in solubility of ACMC as a function of the Mg content is attributed to the increasing water content in the structure of ACMC. A similar behavior has been observed for the hydrated crystalline Mg-carbonate system, where the more hydrated lansfordite ($MgCO_3{\cdot}5H_2O$) exhibits a higher solubility than the less hydrated nesquehonite ($MgCO_3{\cdot}3H_2O$) at 25 °C.[49]

Conclusions

The experimental results indicated a fast exchange of ions between the synthesized amorphous ACMC solids and the aqueous solutions, which is likely stimulated by the large surface area of the nano-porous ACMC solid. Chemical steady state conditions between ACMC and corresponding solution were reached within 2 min of reaction time. The calculated ion activity products documented that the solubility of ACMC increases as a function of the Mg content. This feature is probably linked to the increasing water content and changes in short-range order, as Ca is substituted by Mg in the ACMC structure. Overall, the results reveal a high sensitivity of the ACMC phase in respect to the chemical composition (*e.g.* magnesium concentration) of the aqueous environment. The exchange of ions between the amorphous solid and the solution causes changes in the chemical composition of the solution (*e.g.* pH and alkalinity) and that of the solid (*e.g.* Mg content). This ion exchange probably affects the temporal stability of the amorphous precursor in solution and its subsequent transformation into distinct crystalline calcium magnesium carbonates. Essentially, our findings suggest that the Mg content of ACMC is not the governing parameter for its temporal stability in solution. In contrast, the chemical composition of both the solution and the amorphous solid, have to be taken into account as a potential parameter controlling ACMC stability and transformation behavior. For example, in a closed system the chemistry of the fluid (*e.g.* calcifying fluid in the biomineralisation space) would influence the temporal stability of the amorphous precursor and its transformation pathway to the final crystalline carbonate mineral.

Conflicts of interest

There are no conflicts to declare.

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