

Article

Proton and Electron Irradiations of CH₄:H₂O Mixed Ices

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Abstract: The organic chemistry occurring in interstellar environments may lead to the production of complex molecules that are relevant to the emergence of life. Therefore, in order to understand the origins of life itself, it is necessary to probe the chemistry of carbon-bearing molecules under conditions that simulate interstellar space. Several of these regions, such as dense molecular cores, are exposed to ionizing radiation in the form of galactic cosmic rays, which may act as an important driver of molecular destruction and synthesis. In this paper, we report the results of a comparative and systematic study of the irradiation of CH₄:H₂O ice mixtures by 1 MeV protons and 2 keV electrons at 20 K. We demonstrate that our irradiations result in the formation of a number of new products, including both simple and complex daughter molecules such as C₂H₆, C₃H₈, C₂H₂, CH₃OH, CO, CO₂, and probably also H₂CO. A comparison of the different irradiation regimes has also revealed that proton irradiation resulted in a greater abundance of radiolytic daughter molecules compared to electron irradiation, despite a lower radiation dose having been administered. These results are important in the context of the radiation astrochemistry occurring within the molecular cores of dense interstellar clouds, as well as on outer Solar System objects.

Keywords: astrochemistry; molecular astrophysics; radiation chemistry; methane ice; water ice



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1. Introduction

A comprehensive understanding of the chemistry of carbon-bearing molecules in space is necessary in order to quantitatively study the extent of organic chemistry in extra-terrestrial environments, including reactions that may lead to the formation of complex organic molecules of inherent interest to biology and biochemistry. A number of carbon-bearing molecules have been detected in interstellar space, ranging in structural complexity from methylidyne radicals, CH, to fullerenes such as buckminsterfullerene, C₆₀, and rugby-ballene, C₇₀ [1]. Methane, CH₄, is the simplest saturated organic molecule that is known to exist in the interstellar medium, having originally been observed in both the solid and gas phases towards the young stellar object NGC 7538 IRS 9 [2].

The initial observations of interstellar CH₄ found that a large fraction was locked away in interstellar icy grain mantles, thus suggesting a synthesis pathway that relied upon reactions in the solid phase. Indeed, recent experimental studies have demonstrated that carbon and hydrogen atom addition reactions on the surfaces of interstellar dust grains are able to produce CH₄, and that the rate of formation is twice as high when CH₄ is formed in an ice that is rich in solid water, H₂O [3]. Such experimental results complement the previously observed correlations between CH₄ and H₂O ices in interstellar icy grain

mantles observed towards low-mass young stellar objects [4]. CH₄ has also been observed in comets [5–7], implying the possible survival of interstellar CH₄ during the processes of stellar and planetary system formation. Cometary impacts could indeed have played a key role in the delivery of CH₄ to various planetary-like bodies in the Solar System, such as Titan or Pluto [8,9].

Interstellar and outer Solar System environments are characterized by chemistry driven by ionizing radiation mediated by galactic cosmic rays or the solar wind. As such, it is conceivable that the radiation chemistry of CH₄ in interstellar or outer Solar System ices could be a driver towards the formation of complex organic molecules in these environments. Such radiation chemistry is known to be driven by the release of several thousand low-energy secondary electrons along the track of the irradiating particle [10,11]. Previous laboratory studies have investigated the radiation chemistry of pure CH₄ astrophysical ice analogues using keV-MeV ions and electrons [12–16], and have demonstrated the formation of various daughter hydrocarbons, including ethane, C₂H₆; propane, C₃H₈; ethene, C₂H₄; and ethyne, C₂H₂, as well as several radical species sourced from their mono-dehydrogenation. When mixed with other molecules, irradiated CH₄-containing ices yield even more complex molecules: electron-irradiated mixtures with phosphine, PH₃, yield methylphosphanes as large as CH₃P₈H₉ [17], while the formation of higher-order carboxylic acids such as decanoic acid, C₉H₁₉COOH, has been reported in irradiated mixtures with carbon dioxide, CO₂ [18]. The inclusion of nitrogen-bearing molecules, such as ammonia, NH₃, into electron-irradiated CH₄:CO₂ mixed ices has also been shown to result in the formation of glycine [19], a proteinogenic amino acid of direct relevance to biology.

Of interest to this present study is the radiation chemistry occurring in CH₄:H₂O mixed ices due to the ubiquity of H₂O ices within astronomical environments [20], as well as the apparent association of CH₄ formation with H₂O-rich environments [2,3]. This radiation chemistry has been considered by previous experimental studies [21–24], which have shown that alcohols, higher hydrocarbons, and methanal, H₂CO, are among the readily produced closed-shell daughter molecules. For instance, the study of Wada et al. [22] considered the low-energy electron irradiation of CH₄:H₂O ices at 10 K and found that methanol, CH₃OH, is the major daughter molecule of this process and is formed by the combination of hydroxyl, OH, and methyl, CH₃, radicals, or alternatively by the insertion reaction between methylene radicals, CH₂ and H₂O. In their study on the irradiation of CH₄:H₂O ices by 40 MeV nickel ions at 15 K, Mejía et al. [23] were able to additionally demonstrate the formation of ethanol, C₂H₅OH; ethanal, CH₃CHO; and methanoic acid, HCOOH, along with simpler molecules such as carbon monoxide, CO; CO₂, and H₂CO. Similar results were reported by Krim and Jonusas [24], who also reported the formation of a number of ether molecules, as well as glycolaldehyde, HOCH₂CHO, after the irradiation of a CH₄:H₂O ice by vacuum-ultraviolet photons at 3 K.

To the best of our knowledge, however, no previous study has comparatively investigated the radiation chemistry induced in CH₄:H₂O ices by different types of energetic processing, despite the evident astrochemical importance of such ices. Understanding the potential differences in the radiation chemistries induced by different processing types is important since different astronomical environments are subject to different combinations of ion, electron, and ultraviolet photon irradiations and the predominance of one type of energetic processing over another may lead to the production of distinctive daughter molecules in varying abundances. Indeed, the work of Moore and Hudson [25] demonstrated that several daughter molecules produced after the proton irradiation of CH₄ ices mixed with either CO or molecular nitrogen, N₂, are not produced after similar irradiations using ultraviolet photons. Baratta et al. [12] also performed a comparative study in which they investigated the irradiation of pure CH₄ ices by 30 keV He⁺ and 10.2 eV photons and found that similar administered doses resulted in a greater extent of CH₄ destruction and daughter molecule formation in the case of the He⁺ irradiation. Indeed, they observed that C₂H₆, itself formed from the dissociation of the CH₄ parent species, began to decay after a

dose of approximately 30 eV per 16 u had been administered via He⁺ irradiation; such an observation was not made during irradiation with 10.2 eV photons.

In this paper, we present the results of a systematic and comparative study looking into the irradiation of CH₄:H₂O mixed ices using 1 MeV protons and 2 keV electrons so as to characterize the differences in daughter molecule formation and abundances between the different types of energetic processing. The experimental methodology used to carry out this study is described in Section 2, while our results are presented in Section 3 and are discussed in the context of their applicability to astrochemistry. A summative conclusion is given in Section 4.

2. Methodology

The experimental work was carried out using the Ice Chamber for Astrophysics–Astrochemistry (ICA) at the Institute for Nuclear Research (Atomki) in Debrecen, Hungary. This set-up has been described in great detail in previous publications [26,27] and so, in this paper, we limit ourselves to only the most salient details. The ICA is an ultrahigh-vacuum compatible steel chamber with a base pressure of a few 10^{−9} mbar, which is achieved through the combined action of a turbomolecular pump and a dry rough pump. In the center of the chamber is a gold-coated oxygen-free high-conductivity copper sample holder into which a maximum of four zinc selenide, ZnSe, substrates may be mounted (Figure 1). The sample holder and the substrates may be cooled to 20 K through contact with the cold finger of a closed-cycle helium cryostat, although an operational temperature range of 20–300 K is available.

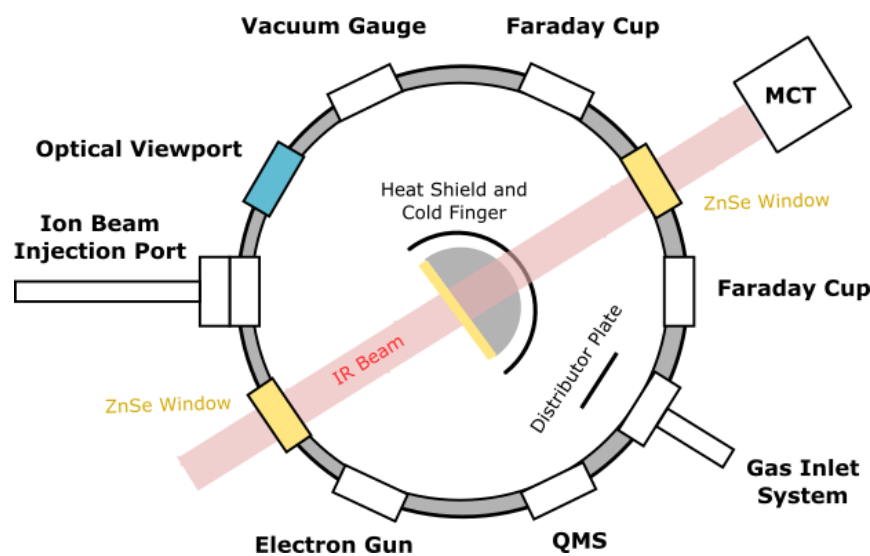


Figure 1. Top-view schematic diagram of the ICA set-up. Figure reproduced from Mifsud et al. [27] with the kind permission of the European Physical Journal (EPJ).

CH₄:H₂O astrophysical ice analogues were prepared on the ZnSe substrates by background deposition of a gas mixture at 20 K. The gas mixture was first prepared in a mixing line with the desired stoichiometric ratio (CH₄:H₂O = 5:2) by introducing CH₄ gas (Linde; 99.5% purity) and H₂O vapor (from a high-purity liquid sample that had been de-gassed by performing several freeze–pump–thaw cycles using liquid nitrogen). The gas mixture was then dosed into the main chamber at a pressure of a few 10^{−6} mbar until ices of approximately 0.25 μm thickness were prepared. The deposition and growth of the ices were monitored in situ using Fourier-transform mid-infrared transmission absorption spectroscopy (resolution = 1 cm^{−1}; spectral range = 4000–650 cm^{−1}) and an external MCT detector. Once deposited, a pre-irradiation mid-infrared spectrum was acquired, which

allowed us to quantify the column densities (N , molecules cm^{-2}) of the deposited molecules via Equation (1):

$$N = \frac{P \ln(10)}{A_v} \quad (1)$$

where P is the area of a characteristic absorption band (cm^{-1}) of the species of interest and A_v is its associated integrated band strength constant (cm molecule^{-1}). The factor of $\ln(10)$ allows for the relation of P (which is measured from spectra on an absorbance scale) to A_v (which is measured on an optical depth scale). Deposited $\text{CH}_4:\text{H}_2\text{O}$ ices were then exposed either to a 1 MeV proton beam delivered by a Tandatron particle accelerator [28,29] or to a 2 keV electron beam delivered by an electron gun mounted directly to one of the side ports of the ICA [27]. The homogeneities, profiles, and currents of the proton and electron beams were respectively assessed using the methods described in Herczku et al. [26] and Mifsud et al. [27] prior to commencing the irradiation of the ices. Beam currents were measured to be 290 nA in the case of the proton beam and 5.7 μA in the case of the electron beam. Additional mid-infrared spectra were collected throughout the irradiation process so as to identify the formation of daughter molecules. A summary of the experiments conducted in this study (including information on the physical parameters required to perform quantitative analysis) is provided in Table 1.

Table 1. Summary of the experiments performed in this study.

Parameter	Experiments	
	1	2
CH_4 column density (10^{17} cm^{-2}) ^a	3.24	3.91
H_2O column density (10^{17} cm^{-2}) ^b	1.29	1.59
Projectile	1 MeV proton	2 keV electron
Projectile stopping power ($\text{eV } \text{\AA}^{-1}$) ^c	2.1	0.9
Fluence delivered (cm^{-2})	4.95×10^{15}	4.16×10^{16}
Dose delivered (eV per 16 u)	46.30	168.42
Irradiation temperature (K)	20	20
Ice thickness (μm)	0.22	0.27

^a Calculated using the area of the ν_4 mode at $\sim 1300 \text{ cm}^{-1}$ assuming $A_v = 1.03 \times 10^{-17} \text{ cm molecule}^{-1}$ [30].

^b Calculated using the area of the ν_s mode at $\sim 3370 \text{ cm}^{-1}$ assuming $A_v = 1.50 \times 10^{-16} \text{ cm molecule}^{-1}$ [31].

^c Calculated using the SRIM software [32].

3. Results and Discussion

The mid-infrared spectra of the $\text{CH}_4:\text{H}_2\text{O}$ ices after their preparation at 20 K, as well as during their processing by 1 MeV proton and 2 keV electron irradiation, are given in Figure 2. Upon preparation, various mid-infrared absorption features are immediately apparent, such as the ν_4 (1300 cm^{-1}), $\nu_2 + \nu_4$ (2816 cm^{-1}), ν_1 (2901 cm^{-1}), and ν_3 (3011 cm^{-1}) modes of CH_4 [16], as well as the ν_L (713 cm^{-1}), ν_2 (1634 cm^{-1}), and ν_s (3370 cm^{-1}) modes of H_2O [33]. Of particular interest is the single, strong absorption band present at around 3616 cm^{-1} , which is attributed to the hydroxyl dangling bond vibration of H_2O [33–36]. In pure H_2O ices, the dangling bond vibration typically presents as two weak absorption features; however, when mixed with CH_4 , these dangling bond absorption bands merge into one significantly stronger band, which is red-shifted to lower wavenumbers. This phenomenon has been studied in detail by Gálvez et al. [37] and Herrero et al. [38], who demonstrated that this could be ascribed to the interaction of CH_4 with microporous H_2O and that the intensity of the band increases with increasing CH_4 content in the mixed ice. Given that our ice is rich in CH_4 ($\text{CH}_4:\text{H}_2\text{O} = 5:2$), it is therefore not surprising that the hydroxyl dangling bond absorption band is one of the strongest features in the acquired mid-infrared spectra (Figure 2). The interaction of CH_4 with microporous H_2O is also responsible for the appearance of the symmetry-forbidden ν_1 mode of CH_4 , which is present in our spectra due to the distortion of the cubic structure of small crystallites, resulting in the breakdown of the symmetry of the CH_4 molecule [38].

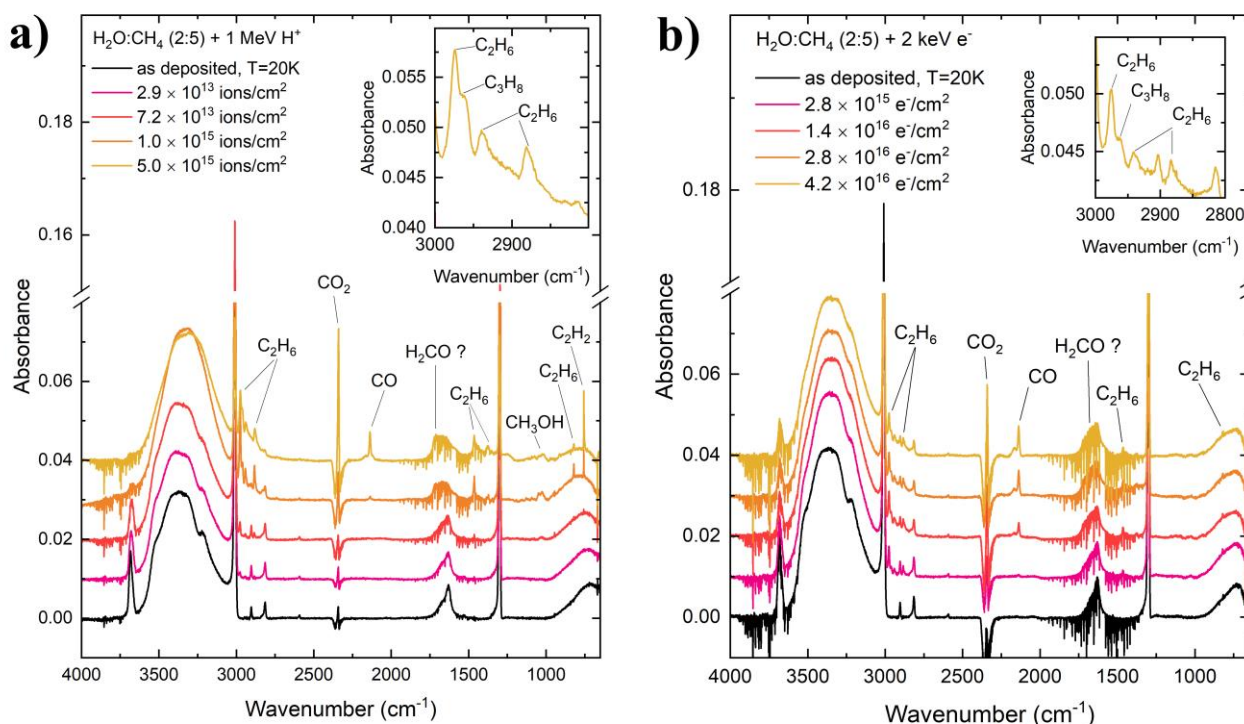
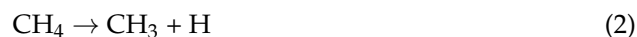
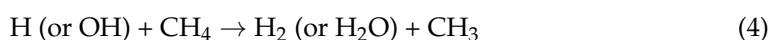


Figure 2. Mid-infrared absorption spectra of CH₄:H₂O (= 5:2) mixed ices upon preparation at 20 K and during irradiative processing using (a) 1 MeV protons and (b) 2 keV electrons. The absorption bands of various daughter molecules have been labelled. Note that the fine structures and negative absorptions in the regions of the H₂O and CO₂ bands are due to variations in the concentrations of atmospheric gas-phase H₂O and CO₂ in the pathlength leading to the detector, which is not maintained under vacuum.

The processing of CH₄:H₂O mixed ices results in the appearance of a number of new absorption bands due to the formation of various daughter molecules as a result of the radiation chemistry occurring within the ices. The most likely first steps leading to the formation of daughter molecules is the radiation-induced dehydrogenation of CH₄ and H₂O to respectively yield CH₃ and OH radicals [21–24].



Hydrogen atoms and OH radicals are likely to be very mobile within the ice structure and may thus partake in various hydrogenation or hydrogen abstraction reactions, including the further dissociation of CH₄ to CH₃ radicals, or the combination reactions leading to the formation of molecular hydrogen, H₂, and hydrogen peroxide, H₂O₂. Neither of these latter two species were detected in our experiments. The non-detection of H₂ is ascribed to its lack of a permanent dipole moment, which makes it infrared-inactive, as well as its volatility, which allows it to efficiently escape from the ice via sublimation to the gas phase; meanwhile, the non-detection of H₂O₂ is likely due to its mid-infrared absorption bands being obscured by those of hydrocarbon species, which are more abundant within the ice.

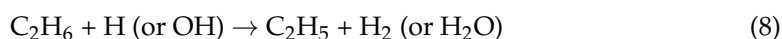


Given the preponderance of CH₄ within our ices, it is most likely that radicals derived from the radiolysis of this species would come into contact with one another. For instance, the combination of two CH₃ radicals would yield C₂H₆ as a result of the formation of a single bond between the two carbon atoms. Indeed, mid-infrared absorption bands attributable to the ν_{12} (822 cm⁻¹), ν_6 (1375 cm⁻¹), ν_{11} (1464 cm⁻¹), ν_5 (2883 cm⁻¹), $\nu_8 + \nu_{11}$ (2941 cm⁻¹), and ν_{10} (2975 cm⁻¹) modes of C₂H₆ [16] were observed after both proton and electron irradiation of the ice mixtures (Figure 2).

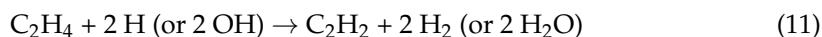
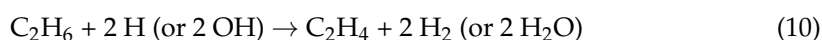


The formation of C₂H₆ is of inherent astrochemical interest since it may be classified as a complex organic molecule (the astrochemical definition of a complex molecule includes all those with six or more constituent atoms). To the best of our knowledge, C₂H₆ has yet to be detected in the interstellar medium [1], although it has been found in various icy outer Solar System objects. Solid C₂H₆ has been observed on the surface of Pluto [39], while liquid C₂H₆ is thought to be an important component of oceanic and lacustrine geochemical systems on Titan [40,41]. Our results conclusively show that C₂H₆ production from CH₄ radiolysis is not inhibited by the presence of H₂O in the ice mixture, and thus it may conceivably form by solid-phase radiation chemistry in realistic interstellar icy grain mantles rich in both species [3,4]. Indeed, observational studies of comet C/1996 B2 Hyakutake revealed an abundance of C₂H₆ comparable to that of CH₄, implying that the comet did not originate in a thermochemically equilibrated astronomical environment [42]. Since the production of C₂H₆ by gas-phase ion–molecule reactions is energetically forbidden [43], this observation strongly points to a solid-phase production route towards C₂H₆ in dense interstellar cloud cores, which is consistent with our experimental results.

The presence of C₂H₆ within the ice opens further radiolytic channels towards the formation of higher hydrocarbons: C₃H₈ is another daughter species formed after the irradiation of our CH₄:H₂O mixed ices (Figure 2). The presence of C₃H₈ in the irradiated mixed ices is likely the result of a two-step process first involving the dehydrogenation of C₂H₆ to yield ethyl radicals, C₂H₅, followed by the addition of CH₃ to yield the observed product.



Furthermore, C₂H₆ may partake in successive dehydrogenation reactions leading to the formation of alkenes and alkynes: C₂H₂ was efficiently formed after the 1 MeV proton irradiation of our CH₄:H₂O mixed ice, as indicated by the intensity of its ν_5 mode at 757 cm⁻¹ [44] (Figure 2). The formation of C₂H₂ from C₂H₆ is a multi-step process involving four dehydrogenation reactions, with C₂H₄ being formed in the process after two hydrogen abstractions. Interestingly, no C₂H₂ was detected after the irradiation of our mixed ice using 2 keV electrons (Figure 2). This observation may be the result of the lower molecular destruction and formation cross-sections associated with projectile electrons compared with protons [45,46], and constitutes a major difference in the irradiation experiments considered in this paper. It is also worth noting that C₂H₄ was not detected in either the proton- or the electron-irradiated ice, although we speculate that this may have been due to the absorption bands of this molecule having been obscured by the stronger absorption bands of the parent CH₄ and H₂O species [16].



Another molecule for which there exists a significant discrepancy in the comparative abundance produced after proton and electron irradiation of the CH₄:H₂O ices is CH₃OH, with more of this molecule being formed after proton irradiation compared to electron irradiation (Figures 2 and 3). The formation of CH₃OH may have occurred as a result of a number of different radiolytic pathways. One such pathway is the direct combination of CH₃ and OH radicals [47].



However, further dehydrogenation of the nascent CH₃ radical may yield other reactive carbon-based radicals, including CH₂ and CH radicals. The reaction between CH₂ and OH, followed by the hydrogenation of the carbon atom, could also yield CH₃OH.



Wada et al. [22] also provided evidence for the insertion of CH₂ into the H₂O molecular structure as an alternative route towards CH₃OH formation.



The necessary precursor radicals in the reactions leading to the formation of CH₃OH highlighted above have all been observed in experimental studies [24], and thus all likely contribute to its formation in our work to at least some extent.

Another formation pathway for CH₃OH in our irradiated ices is the step-wise hydrogenation of CO to sequentially yield HCO, H₂CO, CH₃O (or H₂COH), and finally CH₃OH [48–50]. The contribution of such a reaction sequence to CH₃OH formation in our ices necessitates the presence of CO, which was observed in our irradiated ices. Although the specific mechanism for CO formation as a result of the irradiation of CH₄:H₂O mixed ices is not known for certain, it is possible to suggest that carbon-rich radicals such as CH or even possibly carbon atoms produced as a result of the repeated dehydrogenation of CH₄ may have been able to react directly with H₂O to yield CO. Although speculative, such a suggestion is in agreement with previous works that have demonstrated the formation of both CO and CO₂ as a result of the irradiation of H₂O ice deposited on top of carbonaceous residues and hydrogenated carbon grains [51–53].

It should be noted that the sequential hydrogenation of CO should also yield H₂CO, another closed-shell daughter molecule. This species is usually detected by the observation of its intense ν₂ mode at around 1725 cm⁻¹ [31]. However, in our ices, this detection is made somewhat more complicated by two factors: firstly, this band would at least partially coincide with the broad ν₂ mode of H₂O. Secondly, the appearance of many fine absorption features in this region of the spectrum due to gas-phase H₂O absorption in the pathlength near the spectroscopic detector makes the deconvolution of any possible composite absorption band to its individual contributors practically impossible. However, by looking more closely at Figure 2, it is possible to note that the broad band attributed to the ν₂ mode of H₂O in the pre-irradiation spectra appears to broaden further and blue-shift to higher wavenumbers during irradiation. This could be interpreted as being the result of the appearance of the H₂CO ν₂ mode. However, in the absence of confirmatory spectroscopic deconvolution analyses, we limit ourselves to only tentatively detecting H₂CO in the ices (hence the ‘H₂CO?’ label in Figure 2).

The final molecule to be spectroscopically detected in our irradiation experiments is CO₂ (Figure 2). The formation of CO₂ likely proceeded via a combination of three radiation-driven processes. Firstly, CO₂ formation may occur as a result of the reaction of carbon-rich radicals or carbon atoms with H₂O ice, analogously to the process described previously for CO formation [51–53]. CO₂ is also a known dissociation product of CH₃OH [26,27,54–58], and so the radiolytic destruction of this latter species may contribute somewhat to the formation of CO₂. Lastly, and possibly most significantly, the formation of CO₂ may occur as a result of

the reaction of CO with OH radicals via the production of the HO–CO intermediate [59,60]. Conversely, the reaction between CO and any oxygen atoms yielded from the radiolytic destruction of some oxygen-containing precursor species is not anticipated to contribute to the presence of CO₂ in our ices due to the known high activation energy barriers associated with the reaction [61–63].

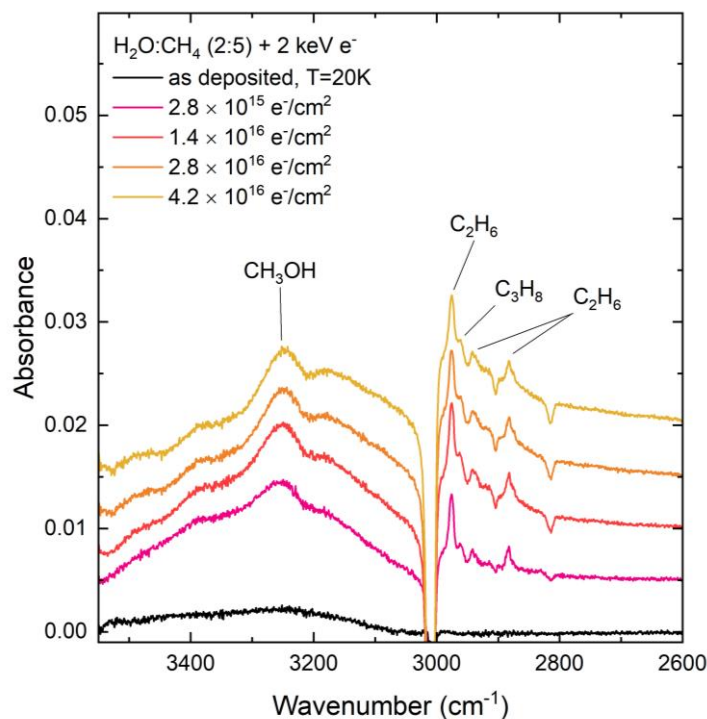
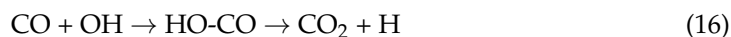


Figure 3. Mid-infrared absorption spectra of a CH₄:H₂O (=5:2) mixed ice upon preparation at 20 K and during irradiative processing using 2 keV electrons. Note that spectra acquired during irradiation are so-called ‘difference spectra’ yielded after the subtraction of the ‘as-deposited’ spectrum. The spectra are focused on the 3500–2600 cm^{−1} region so as to highlight the formation of CH₃OH as a result of the electron irradiation of the CH₄:H₂O mixed ice.

In order to better quantify the differences in the radiation chemistry of our CH₄:H₂O ices mediated by 1 MeV protons or 2 keV electrons, we have compared the column density evolution as a function of the radiation dose for two radiolytic daughter molecules: CO and C₂H₆ (Figure 4). By more closely examining the evolution of the column densities of these molecules in both the proton- and the electron-irradiated ices, it is possible to note that the overall qualitative trend is similar: the column density of C₂H₆ begins to increase until it peaks at a given dose (approximately 19 and 55 eV per 16 u in the case of proton and electron irradiation, respectively), after which it slowly decreases. Conversely, the column density of CO is observed to continually increase with higher radiation doses, with no evidence of peak abundance having yet been reached in either irradiated ice. Interestingly, extrapolation of the proton irradiation data plotted in Figure 4 suggests that the column densities of the C₂H₆ and CO daughter molecules are equal at approximately the same dose (around 63 eV per 16 u) in both the proton- and the electron-irradiated ice.

More generally, it is possible to note that the column densities of daughter molecules measured after proton irradiation are significantly greater than those measured after electron irradiation (Figure 4), despite similar initial (i.e., pre-irradiative) CH₄ and H₂O column densities in each case. Once again, we speculate that this is the result of the larger cross-sections associated with molecular formation and destruction induced by proton irradiation compared to electron irradiation [45,46]. We note that this reason was not invoked by

Baratta et al. [12], who investigated the comparative processing of pure CH₄ ice using 30 keV He⁺ and 10.2 eV ultraviolet photons. In their study, Baratta et al. [12] noted that, at low radiation doses of less than 20 eV per 16 u, the decay of CH₄ and the concomitant production of C₂H₆ followed very similar trends in both the He⁺ and the photon irradiation experiments, both qualitatively and quantitatively. At higher doses, however, the decay of CH₄ proved to be more rapid in the case of He⁺ irradiation. Furthermore, the column density of C₂H₆ eventually peaked and subsequently began to decay during He⁺ irradiation, whereas it eventually plateaued during ultraviolet photon irradiation.

The differences observed by Baratta et al. [12] were ascribed to the dependence of the photochemistry induced by the ultraviolet photons on the optical properties of the ice, which would have undoubtedly changed as a result of its changing chemical composition, thus attenuating many of the incident photons at the most heavily processed regions closest to the surface of the ice. This contrasts with our study, where the radiation chemistry induced by both the projectile protons and electrons was independent of the optical properties of the ice, and both projectile types were able to access practically the entirety of the deposited ice. As such, differences in the total cross-sections of each radiation process seem to be the most parsimonious explanation for the observed differences in the column densities of the radiolytic daughter molecules (Figure 4).

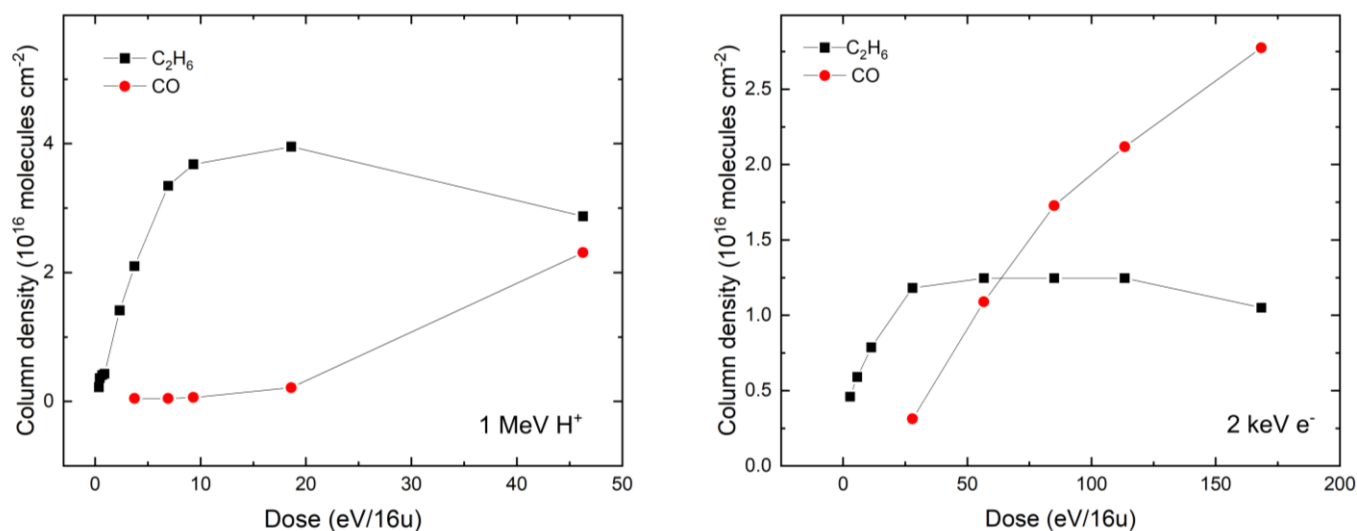


Figure 4. Dose evolution of the column densities of CO and C₂H₆ radiolytic daughter molecules during 1 MeV proton irradiation (left panel) and 2 keV electron irradiation (right panel) of a CH₄:H₂O (= 5:2) mixed ice. Column densities were calculated using integrated band strength constants of 3.51×10^{-18} and 1.12×10^{-17} cm molecule⁻¹ for the C₂H₆ absorption band at 2883 cm⁻¹ and the CO absorption band at 2139 cm⁻¹, respectively [31,64].

4. Conclusions

In this paper, we have reported the results of 1 MeV proton and 2 keV electron irradiations of CH₄:H₂O (= 5:2) mixed ices representative of those that may conceivably exist within icy grain mantles in dense interstellar clouds. Our results have shown that both irradiations result in the formation of many new molecules, some of which are considered to be complex organic molecules in the field of astrochemistry. Among these radiolytic daughter molecules were C₂H₆, C₃H₈, CH₃OH, CO, CO₂, and probably H₂CO. Proton irradiation of the ice mixture additionally resulted in the production of C₂H₂, which was not observed after electron irradiation of the ice, despite a significantly higher radiation dose being administered during electron irradiation. Our results have further demonstrated that the abundances of these daughter species are greater after proton irradiation compared to electron irradiation, possibly due to the larger molecular formation and destruction cross-sections in the case of the former. Interestingly, the column densities of CO and C₂H₆

were found to be equal after an administered radiation dose of approximately 63 eV per 16 u in both the proton- and the electron-irradiated ices.

Our results are of inherent interest to the astrochemistry occurring within icy grain mantles in dense interstellar clouds as they effectively demonstrate the production of complex organic species such as hydrocarbons and alcohols after processing by galactic cosmic rays, for which we have used energetic protons and electrons as an analogue. Such molecules may serve as feedstocks for the production of even more complex molecules of inherent interest to biology and biochemistry. Many of the daughter molecules produced by the irradiation of the CH₄:H₂O ices may also be preserved and seeded to nascent planetary systems, as evidenced by their detections in comets and other Solar System objects. Our results therefore contribute to the better understanding and quantification of daughter molecules that can be produced as a result of different irradiation regimes that may dominate in different astronomical settings.

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