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Three new bismuth(III) pyridine-2,6-dicarboxylate compounds: Synthesis, characterization and crystal structures

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Bismuth, Anionic metal organic compounds, Hydrogen bonding networks, Proton transfer, X-ray crystal structure

abstract

Three new metal-organic compounds containing bismuth and pyridine-2,6-dicarboxylate (pydc) formulated as $(2\text{-apyH})_2[\text{Bi}(\text{pydc})_2(\text{pydcH})] \cdot 2\text{H}_2\text{O}$, **1**, $(4\text{-apyH})[\text{Bi}(\text{pydc})(\text{pydcH})] \cdot 4\text{H}_2\text{O}$, **2** and $(\text{pipzeaH})[\text{Bi}_2(\text{pydc})_3(\text{pydcH})(\text{H}_2\text{O})_2] \cdot 5\text{H}_2\text{O}$, **3**, (2-apy $\frac{1}{4}$ 2-aminopyridine, 4-apy $\frac{1}{4}$ 4-aminopyridine, pipzea $\frac{1}{4}$ 2-piperazin-1-ylethanamine), have been synthesized in deionized water and characterized by elemental analysis (C, H and N), spectral (UV-Vis, IR), ¹H NMR spectroscopy, TGA and single crystal X-ray diffraction. These compounds were obtained via proton transfer methodology. Compounds **1** and **2** have similar monomeric bismuth coordination units, whereas compound **3** has a dimeric bismuth coordination unit. The compounds are anionic in **1** and **2** and they are connected non-covalently to 2-apyH and 4-apyH, respectively. In **3**, two molecules are present, one neutral and one anionic, and both are connected non-covalently to pipzeaH cations. Five different coordination modes of Bi-pydc exist in **1**, **2** and **3**. These compounds are stabilized in the solid state by a complex network of hydrogen bonds between crystal-lization water molecules, anionic, cationic and neutral fragments, forming 3D-supramolecular arrays.

1. Introduction

Design and construction of metal-organic compounds, have been studied for a variety of potential applications [1,7]. Metal-organic compounds containing carboxylic acid ligands in combination with transition metal and lanthanide ions have been studied extensively [8,11], but compounds based on s- and p-block metal ions have obtained popularity only in the recent past [12,18].

Pyridine-2,6-dicarboxylic acid, pydcH₂ is a versatile chelating ligand, with a recognized biological function in the body metabolisms [19], and in a variety of processes such as enzyme inhibitors, plant preservatives [20] and food sanitizers [21]. The complexation of metal ions with pydcH₂ has been the subject of numerous reports [9,10,22,29]. The reasons for this interest are the ability of the ligand to give stable chelates with different coordination modes [22], and its biological activity in human metabolisms [30]. Pyridine-2,6-dicarboxylate complexes have been used as electron carriers in some model biological systems as specific molecular tools in DNA cleavage [31] and as NO scavengers [32].

Bismuth-based materials have many potential applications in the field of photoluminescence, medicine, catalysis and piezoelectric or high dielectric constant materials [33,44]. A large number of bismuth(III) carboxylates are clinically active against Helicobacter pylori and other gastrointestinal disorders [35,44]. The bismuth(III) center is highly acidic and can exhibit variable coordination numbers, framework flexibilities, and different structural characteristics [16,18,34,51]. Bismuth(III) is known to form stable chelate complexes with pydcH₂ (or its anions) and various crystal structures have been reported [16,17,45,51].

In order to gain more information about packing features of molecules containing bismuth(III) chelates, we have studied the synthesis, characterization, spectroscopic and crystal structures of bismuth(III)pydc complexes with 2-apyH, 4-apyH, or pipzeaH cations. We succeeded in synthesizing and characterizing three new crystalline compounds: $(2\text{-apyH})_2[\text{Bi}(\text{pydc})_2(\text{pydcH})] \cdot 2\text{H}_2\text{O}$, **1**, $(4\text{-apyH})[\text{Bi}(\text{pydc})(\text{pydcH})] \cdot 4\text{H}_2\text{O}$, **2** and $(\text{pipzeaH})[\text{Bi}_2(\text{pydc})_3(\text{pydcH})(\text{H}_2\text{O})_2] \cdot 5\text{H}_2\text{O}$, **3**. Compounds **1-3** are notable for several reasons. There is no report in the literature of any

$[\text{Bi}(\text{pydc})_2(\text{pydcH})]^{2-}$, $[\text{Bi}(\text{pydc})(\text{pydcH})]^{-}$, $[\text{Bi}_2(\text{pydc})_4(\text{H}_2\text{O})_2]^{2-}$ and $[\text{Bi}_2(\text{pydc})_2(\text{pydcH})_2(\text{H}_2\text{O})_2]$ complexes with 2-apyH^b, 4-apyH^b and pipzeaH^b as counter ions. Moreover, four different coordination modes of metal-dipicolinate (Scheme 1) exist in **1**, **2** and **3**, a rare coordination type in the bismuth dipicolinate complexes, (b) [13], and the other, (c), which has not been reported in the bismuth dipicolinate complexes.

2. Experimental setup

2.1. Reagents and apparatus

Bi(NO₃)₃·5H₂O (Aldrich, 98%), pyridine-2,6-dicarboxylic acid (pydcH₂) (Merck, 97%), 2-apy $\frac{1}{4}$ 2-aminopyridine (Aldrich, 99%), 4-apy $\frac{1}{4}$ 4-aminopyridine (Aldrich, 98%) and pipzea $\frac{1}{4}$ 2-piperazin-1-ylethanamine (Merck, 97%) were used as purchased. Melting points were determined by applying a Barnstead

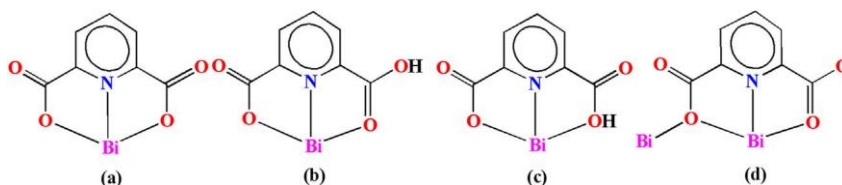
Electrothermal 9200 apparatus. Elemental analyses (C, H, and N) were performed with a Heraeus CHN Pro apparatus. The IR spectroscopy was performed applying a Shimadzu 8400 Fourier transform spectrophotometer ($4000\text{--}4000\text{ cm}^{-1}$) using KBr disc. The NMR spectra were obtained applying a Bruker DRX 250 Avance spectrometer at 2500 MHz using *d*₆-dimethylsulfoxide as solvent, chemical shifts are reported on the δ scale relative to TMS (Scheme S1 and Figs. S1, S3). Thermogravimetric analysis (TGA) experiments were performed on a Netzsch TG 209 instrument at a heating rate of 10 C min^{-1} . Absorption spectra were recorded on a Shimadzu Model 160-A UVeVis spectrophotometer with a 1-cm quartz cell.

Suitable crystals of all the compounds were selected for crystal structure determination using single-crystal XRD. For 1 and 2, unit cell determination and full data collection were carried out on a Bruker APEX II diffractometer [52,53], while for compound 3, a Gemini R Ultra diffractometer [54] has been used, both equipped with a CCD area detector and a graphite monochromatized Mo-K α radiation source. The multi-scan absorption correction has been applied for all the compounds. Data collection, reduction and absorption correction were performed by the Bruker APEX II [53] software for complexes 1 and 2 and with CrisAlisPro [55] software for complex 3.

All structures were solved by direct methods and refined by full-matrix least-square on F^2 using SHELX-97 package [56]. During structure refinements, all atoms (except H) have been refined anisotropically. For compound 1, 2 and 3, the hydrogen atoms of NH and NH₂ groups as well as the water molecules and H(O) were found in difference Fourier synthesis while the H(C) atom positions were calculated. All hydrogen atoms were refined with a riding model and $U_{\text{iso}}(\text{H})$ parameters equal to $1.2U_{\text{eq}}(\text{C}_i)$, $1.2U_{\text{eq}}(\text{N}_i)$ and 1.2 or $1.5U_{\text{eq}}(\text{O}_i)$ where $U_{\text{eq}}(\text{C}_i)$, $U_{\text{eq}}(\text{N}_i)$, $U_{\text{eq}}(\text{O}_i)$ are respectively the equivalent thermal parameters of the carbon, nitrogen and oxygen atoms to which corresponding H atoms are bonded.

The molecular and crystal packing diagrams were generated using Mercury [57]. Geometrical calculations were done using PARST95 [58] present in the WinGX program suite [59].

Further details about data collections and refinements are reported in Table 1.



Scheme 1. Different coordination modes of pydc observed in 1, 2 and 3, where (b) and (c) coordination modes were observed for the first time in Bi^{III} complexes with dipicolinate ligand.

2.2. Synthesis of (2-apyH)₂[Bi(pydc)₂(pydcH)]·2H₂O, 1, (4-apyH) [Bi(pydc) (pydcH)₂]·4H₂O, 2 and (pipzeaH)[Bi₂(pydc) 3(pydcH) (H₂O)₂]·5H₂O, 3

Compound 1 and 3 were synthesized by the refluxing of pydcH₂ (0.33 mmol, 54 mg) and Bi(NO₃)₃·5H₂O (0.11 mmol, 53.36 mg) with different organic amines (2-apy $\frac{1}{4}$ 0.33 mmol, 31 mg; 4-apy $\frac{1}{4}$ 0.33 mmol, 31 mg and pipzea $\frac{1}{4}$ 0.33 mmol, 43 mg) in H₂O solution for 6 h. The resulting solution was cooled down to room temperature and X-ray quality single crystals were obtained by slow solvent evaporation. A synthesis procedure is shown in Scheme 2. For 1: yield of approximate 55% based on bismuth. M.p.: decomposed >255 C. Elemental analysis based on the formula C₃₁H₂₈N₇O₁₄Bi, 1(%): Calc. C, 39.93; H, 3.01; N, 10.52. Found(%): C, 40.01; H, 2.95; N, 10.39. FTIR (KBr, cm⁻¹): 3340(m), 3193(m), 3093(m), 2900(m), 2638(m), 1643(s), 1554 (s), 1527(s), 1434(m), 1380(s), 1272(s), 1195(s), 1072(s), 1018(s), 918(s), 833(s), 732(s), 663(w), 516(w). ¹H NMR (DMSO) $\frac{1}{4}$ 8.32 (s, 3H₄, [pydc]), 8.23 (s, 6H_{3,5}, [pydc]), 7.95 (m, 2H₆, [2-apyH^b]), 7.86 (m, 2H₄, [2-apyH^b]), 6.93 (m, 2H₅, [2-apyH^b]), 6.82 (m, 2H₃, [2-apyH^b]), 3.44 (br, 4H_{Amin}, [2-apyH^b], 2H_{water}) ppm. UVeVis (aqueous solution) (λ , nm): 275. For 2 Yield of approximate 62% based on bismuth. M.p.: decomposed >260 C; elemental analysis based on the formula C₂₆H₂₆N₅O₁₆Bi, 2(%): Calc. C, 37.25; H, 2.98; N, 8.01; Found(%): C, 37.05; H, 2.94; N, 7.96. FTIR (KBr pellet, cm⁻¹): 3340(m), 3155(m), 3050(m), 2910(m), 2692(m), 1620(m), 1535(s), 1427(m), 1373(w), 1265(m), 1188(s), 1072(s), 1010(m), 902(w), 833(w), 763(s), 709(w), 649(w), 509(w). ¹H NMR (DMSO) $\frac{1}{4}$ 8.32 (s, 3H₄, [pydc]), 8.21 (s, 6H_{3,5}, [pydc]), 8.13 (m, 2H₂, [4-apyH^b]), 6.81 (m, 2H₃, [4-apyH^b]), 3.43 (br, 2H_{Amin}, [4-apyH^b], 4H_{water}) ppm. UVeVis (aqueous solution) (λ , nm): 230, 290. For 3 Yield of approximate 52% based on bismuth. M.p.: decomposed >320 C; elemental analysis based on the formula C₃₄H₄₃N₇O₂₃Bi₂, 3(%): Calc. C, 30.55; H, 3.22; N, 7.34; Found(%): C, 30.15; H, 3.24; N, 7.28. FTIR (KBr pellet, cm⁻¹): 3432(m), 3159(m), 3077(m), 2816(m), 1621(s), 1607(s), 1581(s), 1565(s), 1429(s), 1374(s), 1280(m), 1080(s), 1005(m), 911(w), 863(w), 780(s), 735(w), 670(w), 584(w), 529(w). ¹H NMR (DMSO) $\frac{1}{4}$ 8.48 (m, 4H₄, [pydc]), 8.31 (s, 8H_{3,5}, [pydc]), 2.55 (s, 8H_{2,3,5,6}, 4H_{1,2}, [pipzeaH^b]) ppm. UVeVis (aqueous solution) (λ , nm): 270.

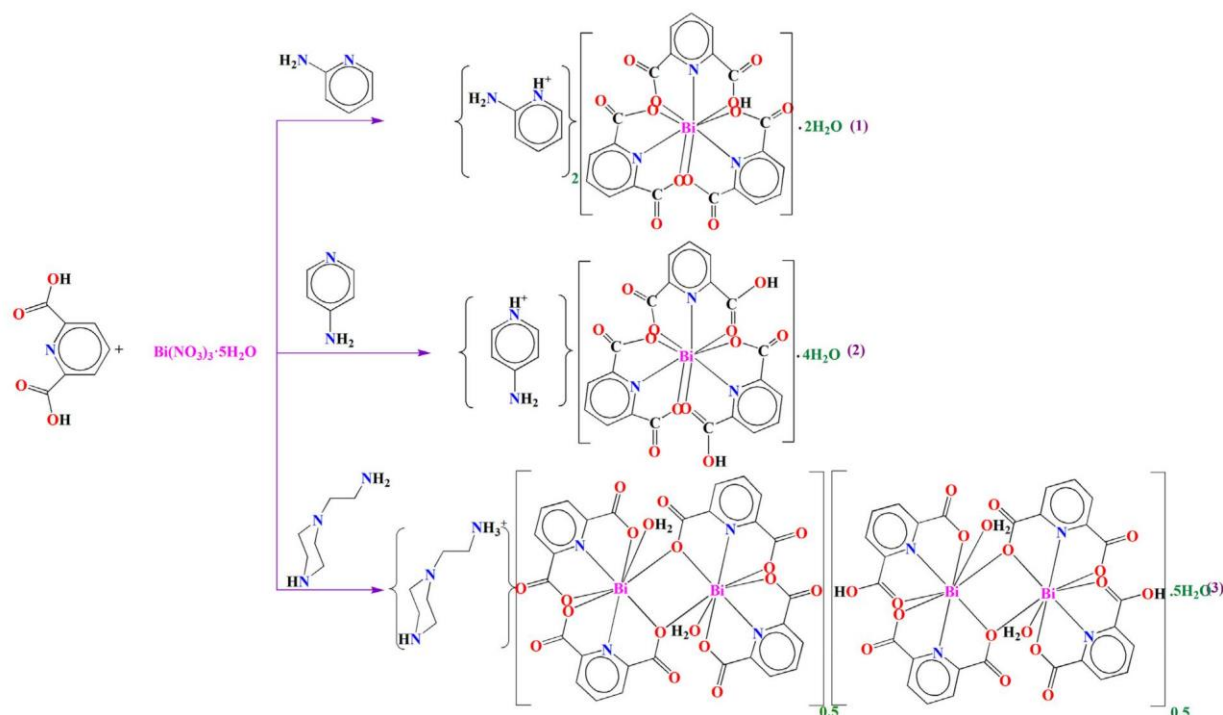
3. Results and discussion

3.1. Spectroscopic studies

Infrared spectroscopy data confirm the coordination of pydc ligand to bismuth (III) ion via the carboxylate group in 1e3. The band associated to the antisymmetric stretching vibrational mode, $\nu_{\text{as}}(\text{eCOO}^-)$, appears at 1643 cm^{-1} and 1620 cm^{-1} in 1 and 2, (1699 cm^{-1} in the free ligand [49]), respectively; while the band associated to symmetric stretching vibrational mode, $\nu_{\text{s}}(\text{eCOO}^-)$, appears at 1380 and 1373 cm^{-1} in 1 and 2 (1332 cm^{-1} in the free

Table 1
Crystallographic data for compounds 1, 2 and 3.

	(2-apyH) ₂ [Bi(pydc) ₂ (pydcH)]·2H ₂ O, 1	(4-apyH)[Bi(pydc)(pydcH) ₂]·4H ₂ O, 2	(pipzeaH)[Bi ₂ (pydc) ₄ (H ₂ O) ₂]·5H ₂ O, 3
Empirical formula	C ₃₁ H ₂₈ N ₇ O ₁₄ Bi	C ₂₆ H ₂₆ N ₅ O ₁₆ Bi	C ₃₄ H ₄₂ N ₇ O ₂₃ Bi ₂
Formula weight	931.58	873.50	1334.71
Temperature/K	294(2)	293(2)	293(2)
Wavelength/Å	0.71073	0.71073	0.71073
Crystal system	Triclinic	Triclinic	Monoclinic
Space group	P 1	P 1	P2 ₁ /c
Z	2	2	4
Unit cell dimensions	a = 11.6659 (5) Å b = 12.0081 (5) Å c = 13.8719 (5) Å α = 94.511 (2) β = 108.139 (2) γ = 111.158 (2)	A = 10.7136(4) Å b = 13.0245(5) Å c = 16.1537(9) Å α = 107.589(1) β = 104.616(1) γ = 103.575(2)	a = 16.3203(4) Å b = 11.9924(3) Å c = 23.1712(6) Å β = 106.757(3)
Absorption coefficient/	5.324 mm ⁻¹	6.05 mm ⁻¹	8.19 mm ⁻¹
Min. and max. transmission factor	0.85e1.00	0.40e1.00	0.28e1.00
F(000)	916	856	2584
Theta range for data collection	1.58 to 35.25	3.23 to 32.68	3.23 to 32.75
Index ranges	18 < h < 18, 19 < k < 19, 21 < l < 22	11 < h < 11, 15 < k < 15, 16 < l < 16	22 < h < 24, 18 < k < 16, 34 < l < 34
Reflections collected	71564	19750	14510
Completeness to theta	99.1% (to theta 35.25)	90.6% (to theta 32.75)	90.6% (to theta 35.75)
Refinement method	Full-matrix least-squares on F	Full-matrix least-squares on F	Full-matrix least-squares on F
Data/restraints/parameters	14969/0/495	6041/8/451	14510/0/585
Goodness-of-fit on F ²	1.034	1.045	1.008
Final R indices [I > 2σ(I)]	R1 = 0.0198, wR2 = 0.0425	R1 = 0.0187, wR2 = 0.0582	R1 = 0.0481, wR2 = 0.0738
R indices (all data)	R1 = 0.0256, wR2 = 0.0438	R1 = 0.0202, wR2 = 0.0642	R1 = 0.0945, wR2 = 0.0869
Largest diff. peak and hole	0.671 and 0.764 e.Å ⁻³	0.878 and 0.639e.Å ⁻³	1.20 and 1.02 e.Å ⁻³



Scheme 2. Preparation route to compounds 1, 2 and 3.

ligand [60]. The value of $D(\nu_{as}(\text{eCOO}^e)\nu_s(\text{eCOO}^e))$, amounts to 263 cm^{-1} for 1, and 247 cm^{-1} for 2, indicating the presence of a carboxylate group coordinated to bismuth(III) ion in the unidentate mode [61,62]. The separation for unidentate carboxylate groups is $>200 \text{ cm}^{-1}$, whereas it is $<200 \text{ cm}^{-1}$ in the bidentate ones [61]. Confirming the presence of $(\text{NH})^b$ group, can be achieved from the presence of bands at 3193 cm^{-1} in 1 as well as 3155 cm^{-1} in 2, due to stretching. Broad band in 3093 cm^{-1} for 1 and 3050 cm^{-1} for 2 is due to stretching vibrations of aromatic CeH's of 2-apyH^b, 4-apyH^b and pydc ions, respectively. The carboxyl OeH stretching frequency (2900 cm^{-1} in 1 and 2910 cm^{-1} in 2) is also observed. Moreover, the sharp bands in 1554 cm^{-1} and 1434 cm^{-1} in 1 as well as 1535 cm^{-1} and

1427 cm⁻¹ in 2 are attributed to aromatic C=C and bands in 1272 cm⁻¹ and 1072 cm⁻¹ in 1 in addition to 1265 cm⁻¹ and 1072 cm⁻¹ in 2, are related to various C=N groups. Finally, sharp bands in 918 and 732 cm⁻¹ in 1 and 902 and 709 cm⁻¹ in 2 are due to out-of-plane aromatic C-H.

The IR spectrum of 3 shows a broad band in 3100–3400 cm⁻¹ due to ν(OH) stretching vibrations of water molecules and pyridine-2,6-dicarboxylic acid; the broadness of this band indicates the presence of hydrogen bonds involving water molecules in 3. The broad band in 3077 cm⁻¹ is due to stretching vibrations of aromatic C-H's of pydc. Also, the broad and branched peak at 2816–3432 cm⁻¹ is attributed to aliphatic C-H's of pipzearH^b and aromatic C-H's of pydc. Strong peak in 1581 cm⁻¹ is attributed to bending vibration of NH₃^b group. The sharp and relatively strong band in 1080 cm⁻¹ is attributed to C=N stretching vibration. In the region of 1565–1429 cm⁻¹, the sharp and relatively strong peaks found are attributed to the aromatic C=C of pydc ring. Also, some strong peaks in 1621 cm⁻¹, together with a strong peak in 1374 cm⁻¹, are due to ν_{as}(eCOO⁻) and ν_s(eCOO⁻), respectively. The value of D (ν_{as}(eCOO⁻) / ν_s(eCOO⁻)), amounts to 249 cm⁻¹ for 3, indicating the unidentate coordination mode of carboxylate group to bismuth(III) ion, as the carboxylate group is considered to be bridging only if both carboxylate oxygen atoms are coordinated to two different metal ions [60]. In the case of 3, only one carboxylate oxygen atom is coordinated to bismuth(III) ions, with the other being uncoordinated, which is similar to unidentate coordination mode of carboxylate group. The value of D is in agreement with this assumption [61]. Sharp bands in 911 and 735 cm⁻¹ in 3 are due to out-of-plane aromatic C-H.

In ¹H NMR spectra (Scheme S1 and Figs. S1, S3), there are three clear characteristic sets of resonances at 8.21 and 8.48 ppm for hydrogens of pydc fragment [10]. The other resonances are attributed to the hydrogens of different amines. The UV-Vis spectra of aqueous solution of 1, 2 and 3 show only one band, illustrated in Fig. S4. The absorption with a maximum in the range 200–300 nm can be assigned to intraligand π/π* and n/π* (carboxyl C=O) transitions of pydc group. To study the thermal stability of complexes 1–3, thermal gravimetric analysis (TG) was performed in the temperature range of 25–800 °C with the heating rate of 10 °C/min in a N₂ atmosphere (Fig. S5). For 1, The TG curve shows that this compound exhibits three steps of weight loss. The first stages are related to removal of two molecules of lattice water in the range 95–150 °C with a weight loss of 3.78% (theoretical loss 3.86%). The second step between 230 °C and 320 °C is related to the release of the 2-apo with a weight loss of 20.76% (theoretical loss 20.20%). The third stage between 350 °C and 450 °C corresponds to the loss of pydc molecules followed by complete destruction of the structure in 800 °C. The final mass remnant of 54.03% is indicative of deposition of Bi₂O₃, (theoretical residue of Bi₂O₃: 50.02%). For 2, The TG curve shows that this compound exhibits three steps of weight loss. The first stages are related to removal of four molecules of lattice water in the range 80–160 °C with a weight loss of 8.72% (theoretical loss 8.24%). The second step between 260 °C and 330 °C is related to the release of the 4-apo with a weight loss of 10.89% (theoretical loss 10.77%). The third stage between 340 °C and 460 °C corresponds to the loss of pydc molecules followed by complete destruction of the structure in 800 °C. The final mass remnant of 54.03% is indicative of deposition of Bi₂O₃, (theoretical residue of Bi₂O₃: 53.34%). The TG curve for compound 3 shows that this compound exhibits three steps of weight loss. The first stages in the range 80–280 °C are related to removal of crystallization water molecules and coordinated water molecules (found: 9.42%, calcd: 9.44%). The second stage in the range 300–350 °C is related to the release of pipzear (found: 8.29%, calcd: 9.68%). The third stage from 350 °C to 400 °C corresponds to the loss of pydc molecule followed by complete destruction of the structure to 800 °C. Nevertheless the final remnants of 45.55% suggests that the complex does not decompose completely under the experimental temperature (theoretical residue of Bi₂O₃ 34.91%).

3.2. X-ray crystallographic structures

The crystallographic data of 1, 2 and 3 are given in Table 1, selected bond lengths and angles are listed in Table 2 and the hydrogen bond geometry in Table 3.

3.2.1. Structures of (2-apoH)₂[Bi(pydc)₂(pydcH)]·2H₂O, 1 and (4-apoH)[Bi(pydc)(pydcH)₂]·4H₂O, 2

Both compounds 1 and 2 crystallized in a triclinic space group and have monomeric units in which the pydc or pydcH moieties act as tridentate ligands, H-bonded to 2-apo or 4-apo cations. The protonated position in the cationic counter ions of both is the pyridine nitrogen atom. In the anionic complexes of 1 and 2, the Bi(III) is nine-coordinated by three nitrogen atoms and six oxygen atoms of carboxylate from pydc, protonated or deprotonated oxygen atom from pydcH fragments. The Bi-O bond lengths are in the range 2.2637(29)–2.8281(27) Å and the Bi-N bond lengths are in the range 2.4324(15)–2.6665(15) Å (Table 2). The average values of Bi-O and Bi-N bond lengths are comparable with the corresponding ones reported in the literature for related bismuth(III) complexes with pyridine-2,6-dicarboxylate or 4-hydroxy pyridine-2,6-dicarboxylate [12,15–18,35,36,45–51].

Table 2
Selected bond distances and bond angles (Å) for Compounds 1, 2 and 3.

Compound 1, (2-apoH) ₂ [Bi(pydc) ₂ (pydcH)]·2H ₂ O			
Bi1eO10	2.3216(11)	Bi1eO32	2.3702(10)
Bi1eO12	2.5252(14)	Bi1eN10	2.4324(15)
Bi1eO21	2.4948(13)	Bi1eN20	2.5683(12)
Bi1eO22	2.5297(13)	Bi1eN30	2.6665(15)
Bi1eO30	2.7946(10)		
O10eBi1eO32	77.757(41)	O22eBi1eO30	86.463(39)
O32eBi1eO22	81.603(44)	O10eBi1eO21	86.840(48)
O10eBi1eO22	77.837(45)	O12eBi1eO32	91.889(42)
O21eBi1eO30	84.676(42)	N10eBi1eN30	120.392(43)
O12eBi1eO30	69.400(36)	N20eBi1eN30	112.952(38)
O21eBi1eO12	76.144(48)	N10eBi1eN20	125.281(39)
Compound 2, (4-apoH)[Bi(pydc)(pydcH) ₂]·4H ₂ O			
Bi1eO4	2.754 (2)	Bi1eO2	2.264 (2)
Bi1eO5	2.577 (2)	Bi1dN1	2.539 (2)
Bi1eO7	2.397 (2)	Bi2dN2	2.497 (2)
Bi1eO8	2.366 (2)	Bi3dN3	2.636 (2)
Bi1eO9	2.828(3)		
O7eBi1eO8	73.741(88)	O2eBi1eO5	84.287(85)
O2eBi1eO8	79.644(90)	O7eBi1eO9	87.550(78)

O7eBi1eO2	79.864(90)	O4eBi1eO8	81.837(76)
O4eBi1eO9	78.881(78)	N1eBi1eN3	113.604(75)
O5eBi1eO9	82.396(81)	N2eBi1eN3	116.162(78)
O4eBi1eO5	87.484(79)	N1eBi1eN2	129.990(75)

Compound 3, (pipzeaH)[Bi2(pydc)3(pydcH)(H2O)2]·5H2O

Bi1eO1A	2.376 (3)	Bi1eO3B	2.561 (3)
Bi1eO3A	2.466 (3)	Bi1eO3B ⁱ	2.632 (3)
Bi1eN1A	2.439 (4)	Bi1eN1B	2.445 (4)
Bi1eO1B	2.308 (3)	Bi1eO11	2.505 (3)
Bi2eO1C	2.399 (4)	Bi2eO1D	2.626 (3)
Bi2eO3C	2.391 (4)	Bi2eO3D	2.307 (3)
Bi2eN1C	2.425 (4)	Bi2eN1D	2.496 (4)
Bi2eO1D ⁱⁱ	2.564 (3)	Bi2eO21	2.510 (4)

Symmetry codes: (i) x, yp1, zp1; (ii) xp1, y, zp1.

O1BeBi1eO1A	91.8 (1)	O3DdB2dO3C	87.7 (1)
O1AdBi1dO3B	75.75 (12)	O1D ⁱⁱ dBi2dO1D	71.45 (12)
O3AdBi1dO3B ⁱ	90.21 (11)	O3CeBi2eO1D ⁱⁱ	93.05 (13)
N1AdBi1dN1B	125.08 (13)	N1CeBi2eN1D	129.59 (14)
O1AdBi1dO3A	132.1 (1)	O3CeBi2eO1C	132.9 (1)
O1AdBi1dO11	147.5 (1)	O3CeBi2eO21	151.2 (1)
O3AdBi1dO3B	146.3 (1)	O1CeBi2eO1D	140.5 (1)
N1AdBi1dN1B	125.1 (1)	N1CeBi2eN1D	129.6 (1)

Symmetry codes: (i) x, yp1, zp1; (ii) xp1, y, zp1

Table 3
Hydrogen bond geometry.

DeH\$\$\$A	DeH	H/A	D/A	< DeH/A
Compound 1, (2-apyH) ₂ [Bi(pydc) ₂ (pydcH)]·2H ₂ O.				
O1eH1A\$\$\$O13 ⁱ	0.83 (5)	2.04 (5)		2.859 (3) 172 (4)
O1eH1B\$\$\$O13	0.77 (4)	1.99 (4)		2.753 (3) 172 (4)
O2eH2A\$\$\$O33 ⁱⁱ	0.86 (3)	2.04 (3)		2.890 (2) 174 (2)
O2eH2B\$\$\$O12 ⁱⁱⁱ	0.79 (3)	2.04 (3)		2.8221 (18) 175 (2)
O30eH20A\$\$\$O20 ^{iv}	0.81 (2)	1.75 (2)		2.5516 (19) 171 (2)
N70eH70\$\$\$O22	0.9100	1.8300		2.730 (2) 169.00
N71eH71A\$\$\$O23	0.8600	1.9900		2.834 (2) 170.00
N71eH71B\$\$\$O20 ^v	0.8800	2.1200		2.926 (2) 153.00
N80eH80\$\$\$O2 ^v	0.8600	1.9400		2.790 (2) 168.00
N81eH81A\$\$\$O1 ⁱⁱ	0.8800	1.9800		2.823 (3) 159.00
N81eH81B\$\$\$O11	0.8400	1.9800		2.818 (3) 178.00
C15eH15\$\$\$O32 ⁱⁱ	0.9300	2.5000		3.258 (2) 138.00
C34eH34\$\$\$O11 ^{vi}	0.9300	2.4000		3.119 (2) 134.00
C72eH72\$\$\$O33 ^{vii}	0.9300	2.4800		3.265 (3) 142.00
C84eH84\$\$\$O31 ^{viii}	0.9300	2.4800		3.140 (3) 128.00
C85eH85\$\$\$O31 ^{ix}	0.9300	2.3400		3.224 (3) 158.00
Symmetry codes: (i) x, y, zp1; (ii) xp1, yp1, zp1; (iii) x, yp1, z; (iv) x, y, z; (v) xp1, y, z; (vi) x, y 1, z; (vii) xp2, yp1, zp1; (viii) xp1, y, z; (ix) xp1, yp1, z.				
Compound 2, (4-apyH)[Bi(pydc)(pydcH)]·4H ₂ O.				
O3eH3\$\$\$O16	1.00 (3)	1.64 (4)		2.568 (5) 153 (4)
O10eH10A\$\$\$O17	1.01 (4)	1.60 (4)		2.577 (4) 162 (4)
O14eH14A\$\$\$O17 ⁱ	0.85	2.12		2.9059 153
O15eH15A\$\$\$O7	0.95	1.94		2.874 (4) 169
O15eH15B\$\$\$O14	0.91	2.09		2.9722 163
O16eH16A\$\$\$O11 ⁱⁱ	0.85	2.24		2.922 (5) 138
O16eH16B\$\$\$O14 ⁱⁱⁱ	0.85	1.94		2.7594 162
O17eH17A\$\$\$O13 ⁱⁱⁱ	0.81	1.92		2.726 (4) 169
O17eH17B\$\$\$O12 ^{iv}	0.93	1.83		2.756 (4) 169
N4eH4\$\$\$O8	1.03	2.60		3.162 (5) 114
N4eH4\$\$\$O15	1.03	1.87		2.830 (4) 153
N6eH6A\$\$\$O12 ^v	1.01	2.07		2.961 (4) 146
N6eH6B\$\$\$O1 ^{vi}	1.01	1.93		2.894 (4) 158
C1eH1\$\$\$O13 ^{vii}	1.08	2.46		3.409 (4) 145
C2eH2\$\$\$O5 ^v	1.08	2.60		3.552 (4) 146
C19eH19\$\$\$O2 ^{viii}	1.08	2.54		3.187 (4) 117
C25eH25\$\$\$O8	1.08	2.51		3.128 (5) 115
C28eH28\$\$\$O1 ^{vi}	1.08	2.41		3.248 (5) 134
C29eH29\$\$\$O5 ^{ix}	1.08	2.44		3.295 (5) 135
Symmetry codes: (i) xp1, yp2, z; (ii) xp1, yp1, z; (iii) x 1, y, z; (iv) x, yp2, zp1; (v) xp1, yp1, zp1; (vi) xp2, yp1, zp1; (vii) x, y 1, z; (viii) xp1, yp2, zp1; (ix) xp1, y, z.				
Compound 3, (pipzeaH)[Bi ₂ (pydc) ₃ (pydcH)(H ₂ O) ₂]·5H ₂ O				
O11eH11A\$\$\$N2E ⁱⁱ	0.87			2.993 (8) 145
		2.24		
O11eH11B\$\$\$O4D ⁱⁱ	0.87			2.689 (6) 174
		1.82		
N1EdH1EA\$\$\$O2A ⁱⁱⁱ	0.86			2.830 (6) 100
		2.55		
N1EdH1E\$\$\$O4A ^{iv}	0.86			2.831 (6) 159
		2.54		
N3EdH3N\$\$\$O3A ⁱⁱ	0.89			2.872 (6) 153
		2.05		
N3EdH3NB\$\$\$O2W ^v	0.89			2.995 (6) 152
		2.18		
N3EdH3N\$\$\$O5W ⁱⁱ	0.89			2.920 (6) 120
		2.37		
N3EdH3N\$\$\$O4C ⁱⁱ	0.89			2.854 (7) 157
		2.02		
O11eH11A\$\$\$N2E ⁱⁱ	0.87			2.993 (8) 145
		2.24		
O11eH11B\$\$\$O4D ⁱⁱ	0.87			2.689 (6) 174
		1.82		
O21eH21A\$\$\$O2B ^{vi}	0.85			2.685 (6) 149
		1.92		
O21eH21B\$\$\$O2W	0.83			2.789 (6) 174
		1.96		
O2AdH2AA\$\$\$N1E ^(x)	0.84			2.830 (6) 142
		2.12		
O1WeH1WA\$\$\$O3W ^{vii}	0.88			2.875 (6) 103
		2.55		
O1WeH1WB\$\$\$O3W ^{vii}	0.83			2.875 (6) 103
		2.57		

Table 3 (continued)

DeH\$\$\$A	DeH	H/A	D/A	< DeH/A
O2WeH2WA\$\$\$O4W ⁱⁱ	0.85			2.821 (6) 161

O2WeH2WB\$\$\$O2A ⁱⁱⁱ	0.97	2.01	2.847 (6)	157
O3WeH3WA\$\$\$O2C ⁱⁱ	0.84	1.93	2.978 (5)	146
O3WeH3WB\$\$\$O2D ^v	0.96	2.25	2.761 (5)	148
O5WeH5WA\$\$\$O1W	0.96	1.90	2.784 (6)	141
O5WeH5WB\$\$\$O1C ^{vi}	0.77	1.98	2.936 (4)	165
C3BeH3B\$\$\$O2D ^{vi}	0.93	2.18	3.506 (7)	173
C4BeH4BA\$\$\$O1W ^{xi}	0.93	2.58	3.427 (6)	140
C4DdH4D\$\$\$O2C ^{ix}	0.93	2.66	3.028 (6)	121
C5DdH5D\$\$\$O4B ⁱⁱ	0.93	2.45	3.379 (6)	168
C1EdH1EC\$\$\$O3A ⁱⁱ	0.97	2.46	3.245	123
C3EdH3E\$\$\$O4W ⁱⁱ	0.97	2.62	3.134 (7)	120
C3EdH3E\$\$\$O1I ⁱⁱ	0.97	2.54	3.354 (7)	137
C6EdH6EA\$\$\$O5W ⁱⁱ	0.97	2.58	3.209 (6)	120
		2.61		

Symmetry codes: (ii) xp1, yp1, zp1; (iii) xp1, y, z; (iv) xp1, y 1/2, z 1/2; (v) x, yp1, z; (vi) xp1, y, zp1; (vii) xp1, y 1/2, zp1/2; (ix) x, y 1/2, z 3/2; (x) x 1, y, z; (xi) x, yp1/2, zp1/2.

The three oxygen atoms (O30, O12 and O21 in 1 as well as O7, O8 and O2 in 2) form a triangle and the other three (O10, O22 and O32 in 1, also O4, O5 and O9 in 2) form another triangle. Considering the angles between the oxygen atoms in both 1 and 2, it is found that they are almost eclipsed. So a prism consisted of six oxygen atoms with three caps of nitrogen atoms on its faces are proposed for both.

3.2.1.1. (2-apyH)₂[Bi(pydc)₂(pydCH)]·2H₂O, 1. Compound 1, has an asymmetric unit with one Bi³⁺ cation, two pydc dianions, one pydCH anion, two 2-apyH cations and also two water molecules

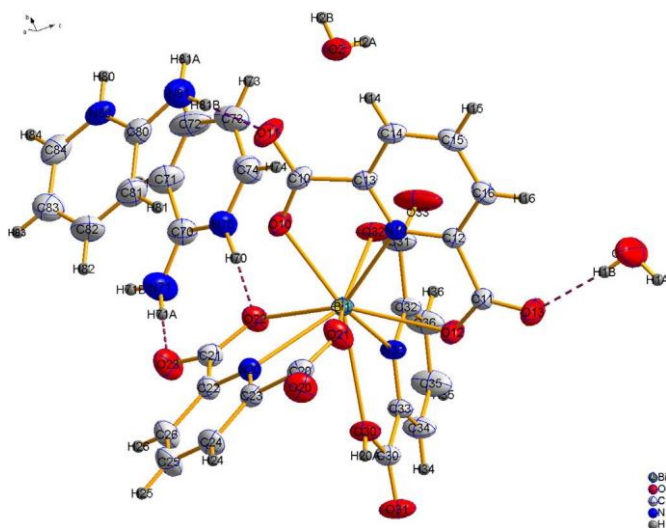


Fig. 1. The molecular and crystal structure of (2-apyH)₂[Bi(pydc)₂(pydCH)]·2H₂O, 1, showing the atom-numbering scheme and displacement. Ellipsoids are at the 50% probability level. Hydrogen bonds are shown with dashed lines.

(Fig. 1). The Bi³⁺ cation is nine-coordinated by six oxygen atoms and three nitrogen atoms from two pydc and one pydcH anions (Fig. S6). The BiN₃O₆ polyhedron is holodirected with a stereo-chemically inactive lone pair of electrons (Fig. S7). The Bi–O bond lengths are in the range 2.3216(11)–2.7946(10) Å and the Bi–N bond lengths are in the range 2.4324(15)–2.6665(15) Å. The bond distance of Bi–O₃₀ (protonated oxygen) is significantly longer (2.7946(10) Å) than the Bi–O_{deprotonated} oxygen bonds (2.3216(11)–2.5297(13) Å). However, it is shorter than the maximum of 3.429 reported for Bi–O retrieved from the Cambridge Structural Database (CSD) [63]. All two pydc and one pydcH anions exhibit similar coordination modes (Fig. S6) and each coordinates only to one Bi³⁺ cation.

The carboxylate groups and the pyridyl nitrogen in two pydc and one pydcH anions chelate in a tridentate manner with the same Bi³⁺ cation to form a [Bi(pydc)₂(pydcH)]²⁻ anionic coordination complex with six uncoordinated carboxylate or carboxylic acid oxygens (two from each pydc or pydcH anion). The cation coordination can be rationalized in terms of bond valence [64e66], which assumes that the total charge of the cations has to be saturated by the summation of the separated bond valence s_i of each coordinated atom or ion. The quantity s_i can be calculated by the expression $s_i = \frac{1}{4} \exp[(r_0 - r_i)/B]$, where r_i is the actual Bi–O or Bi–N distance, r_0 and B are parameters empirically determined ($r_0 = 2.094$ for Bi–O, $r_0 = 2.02$ for Bi–N, $B = 0.37$ for Bi–O and $B = 0.35$ for Bi–N). In the present structure, the value of $\frac{1}{4} s_i$ obtained for the bismuth cation (Table 4), is in good agreement with the expected charge value of 3.00.

Each anionic monomer interacts non-covalently with three 2-

Table 4
Bond Valence Sum (BVS) for Bi atoms in Compounds 1, 2 and 3.

Bond typ	r_i	r_0	B	$r_0 - r_i / B$	B.V
Compound 1, (2-apyH) ₂ [Bi(pydc) ₂ (pydcH)]·2H ₂ O					
Bi–O12	2.525	2.094	0.37	1.1648	0.31198
Bi–N10	2.432	2.02	0.35	1.177	0.30820
Bi–O10	2.323	2.094	0.37	0.6189	0.538
Bi–O21	2.495	2.094	0.37	1.083	0.3385
Bi–N20	2.567	2.02	0.35	1.565	0.2090
Bi–O22	2.530	2.094	0.37	1.178	0.3079
Bi–O30	2.795	2.094	0.37	1.8945	0.150
Bi–N30	2.666	2.02	0.35	1.845	0.158
Bi–O32	2.370	2.094	0.37	0.746	0.474
Compound 2, (4-apyH)[Bi(pydc)(pydcH) ₂]·4H ₂ O					
Bi–e O4	2.754 (2)	2.094	0.37	1.7838	0.1680
Bi–e O5	2.577 (2)	2.094	0.37	1.3054	0.2711
Bi–e O7	2.397 (2)	2.094	0.37	0.8189	0.4409
Bi–e O8	2.366 (2)	2.094	0.37	0.7351	0.4794
Bi–e O9	2.828 (3)	2.094	0.37	1.9838	0.1375
Bi–e O2	2.264 (2)	2.094	0.37	0.4595	0.6316
Bi–1dN1	2.539 (2)	2.02	0.35	1.4829	0.2270
Bi–2dN2	2.497 (2)	2.02	0.35	1.363	0.2559
Bi–3dN3	2.636 (2)	2.02	0.35	1.760	0.1720
Compound 3, (pipzeaH)[Bi ₂ (pydc) ₃ (pydcH)(H ₂ O) ₂]·5H ₂ O					
Bi–e O1B	2.308 (3)	2.094	0.37	0.5784	0.5608
Bi–e O1A	2.376 (3)	2.094	0.37	0.7622	0.4667
Bi–e O3A	2.467 (3)	2.094	0.37	1.0081	0.3649
Bi–e O11	2.504 (3)	2.094	0.37	1.1081	0.3302
Bi–e O3B i	2.561 (3)	2.094	0.37	1.2622	0.2830
Bi–e O3B	2.631 (3)	2.094	0.37	1.4513	0.2342
Bi–e N1A	2.440 (4)	2.02	0.35	1.2	0.3012
Bi–e N1B	2.446 (4)	2.02	0.35	1.2171	0.2961
Bi–2e O3D	2.308 (3)	2.094	0.37	0.5784	0.5608
Bi–2e O3C	2.391 (4)	2.094	0.37	0.5784	0.5608
Bi–2e O1C	2.398 (4)	2.094	0.37	0.8216	0.4397
Bi–2e O21 ii	2.508 (4)	2.094	0.37	1.1189	0.3266
Bi–2e O1D	2.564 (3)	2.094	0.37	1.2703	0.2807
Bi–2e O1D	2.626 (3)	2.094	0.37	1.4378	0.2374
Bi–2e N1C	2.424 (4)	2.02	0.35	1.1543	0.3153
Bi–2e N1D	2.497 (4)	2.02	0.35	1.3628	0.2559

apyH cations through carboxylate-2-apyH H-bonding. Each coordination bismuth unit is connected to two other such units through the 2-apyH cations by H-bonding, leading to the formation of an infinite three-dimensional H-bonded layer structure. As shown in Fig. 2, the carboxyl groups of the pydc and pydcH ligands and water molecules are involved in intermolecular OeH\$\$\$O hydrogen bonding, and form three types of robust hydrogen bond synthons, namely, $R^2_2(8)$ I, $R^2_2(12)$ II and $R^2_2(16)$ III. The other set of hydrogen bonding patterns included NeH/O and CeH/O, are held between the (2-apyH)^b, its counter anion and water molecules (Table 3). An outstanding feature of compound 1 is the presence of CeH/p stacking interactions. The CeH\$\$\$p distance (measured to the centre of the ring) is 2.990 Å (1 x, y, z) and CeH/p angle is 153.42 (Fig. 3). Other intermolecular interactions that can be observed in the structure of compound 1 are pep stackings between aromatic rings, the strongest one is between N70/C70eC74 and N80/C80eC84 rings with the distance of 3.5847(12) Å between the correspondent centers of gravity (Cg), as shown in Fig. 3. The stacking interactions of the carbonyl groups with the ligands are of importance in the structure of 1 explaining the short intermolecular contacts between O(23)/C(21), 3.160 Å, C(21)/C(21), 3.552 Å (Fig. 3). The spaces between two layers of [Bi(pydc)₂(pydcH)]^{2e} fragments are filled with layers of (2-apyH)^b cations and uncoordinated water molecules. The crystal lattice is aggregated through the concert of intermolecular interactions, such as electrostatic attraction between ion pairs, stacking interactions of the carbonyl groups, different kinds of hydrogen bonding (NeH/O, OeH/O and CeH/O), pep and CeH/p stacking interactions.

3.2.1.2. (4-apyH)[Bi(pydc)(pydcH)₂]-4H₂O, 2. Compound 2, has an asymmetric unit with one Bi^{3b} cation, one pydc dianions, two pydcH anion, one 4-apyH cation and also four water molecules (Fig. 4). The Bi^{3b} cation is nine-coordinated by six oxygen atoms and three nitrogen atoms from one pydc and two pydcH anions (as in Fig. S8). Similar to 1 the BiN₃O₆ polyhedron is holodirected with a stereochemically inactive lone pair of electrons. The Bi-eO bond lengths are in the range 2.264(2)e2.828(3) Å and the Bi-eN bond lengths are in the range 2.497(2)e2.636(2) Å. The bond distances of Bi-eO₄ protonated acid (2.754(2) Å) and Bi-eO₉ protonated acid (2.828(3)Å) are slightly longer than the Bi-eO_{deprotonated} oxygen bonds (2.264(2)e 2.577(2) Å). All one pydc and two pydcH anions exhibit similar

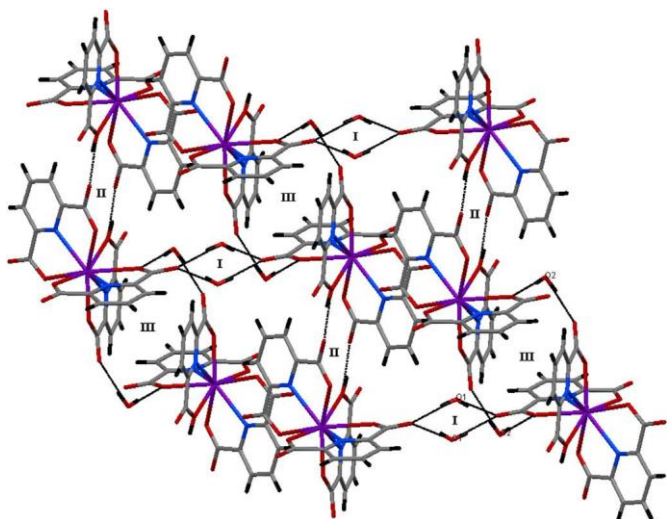


Fig. 2. Graph set descriptors made by OeH\$\$\$O interactions between water ion and [Bi(pydc)₂(pydcH)]^{2e}.

and O17) and O atoms of pydc or pydcH form a cyclic graph set that can be described as $R^3_3(10)$ I, $R^{10}_6(20)$ II, $R^6_4(12)$ III; also, the water molecular cluster with $n \approx 4$ in the crystal structure of **2** consist of a triangle shape formed by three water molecules in which the last water molecule is in the center (Fig. 5a). A base chain, including water molecules, with O...O distances in the range 2.759–2.972 Å is present between the anionic complexes. The other set of hydrogen bonding patterns, including N...O and C...O, are held between the water molecules, (4-apyH)^b and its counter anion (Fig. 5b and c). Repetition of weaker C...O interactions create the hydrogen bonded layer A shown in Fig. 5c.

Each N(4) nitrogen atom of (4-apyH)^b is linked as a bifurcated donor to [Bi(pydc)(pydcH)₂]⁻ unit and water fragment via hydrogen bonding interactions (Table 3 and Fig. 5b). Considering the D_{H/A} distance 2.830 (4) Å for (N4eH4...O15), these intra-molecular interactions are certainly significant. Also the hydrogen bond donor atom in N6eH6A...O12(v) [(v): xp1, yp1, zp1] and N6eH6A...O1(vi) [(vi): xp2, yp1, zp1] is the coordinated N(6), so the donor bridged hydrogen bond occurs.

Other intermolecular interactions that can be observed in the structure of compound **2** are CeO/p stackings, the strongest one is between C7eO1 and N4/C25eC29 (x,y,z) ring with the distance of 3.167(3) Å (measured to the C_g of the ring), shown in Fig. 6.

3.2.2. (pipzeaH)[Bi₂(pydc)₃(pydcH)(H₂O)₂].5H₂O, **3**

The dimeric bismuth pyridine-2,6-dicarboxylate, **3**, has an asymmetric unit of 66 non-hydrogen atoms with two Bi³⁺ cations, three pydc anions and one pydcH anion, one pipzeaH cation, two terminal water molecules (one for each bismuth ion) and 5 free water molecules (Fig. 7a). Applying the inversion center on the asymmetric unit, the molecular structure of the compound is evi-denced (Fig. 7b): one dimeric neutral [Bi₂(pydc)₂(pydcH)₂] complex and one dimeric anionic [Bi₂(pydc)₄]²⁻ complex are formed (Fig. S9), which represents the most common bridging motif in structurally characterized Bi^{III} complexes with dipicolinate ligands [16,36,46,47,51]. Although there are two distinct bismuth atoms in the asymmetric unit of **3**, both Bi³⁺ cations are eightfold coordinated by five oxygen atoms from three different pydc anions, one oxygen from a terminal water molecule, and two nitrogen atoms from two different pydc anions; all the pydc anions chelate in a tridentate manner with the Bi³⁺ cations involving not only the carboxylate oxygens but also the pyridyl nitrogens (Bi-N 2.451(4) Å av.).

Unlike compound **1** and **2**, the bismuth cation in compound **3** is located in a slightly distorted bicapped octahedral coordination environment, BiN₂O₆, in which a small effect of the lone pair on the structure asymmetry is observed (Fig. 7c). For each dimer, two of the four pydc fragments bridge the two bismuth ions, each with one oxygen (O3B and O1D). The bridges are asymmetrical

(Bi1eO3B $\frac{1}{4}$ 2.561(3) Å, Bi1eO3Bⁱ $\frac{1}{4}$ 2.632(3) Å) (i: x, yp1, zp1),

Bi2eO1D $\frac{1}{4}$ 2.626(3) Å, Bi2eO1Dⁱ $\frac{1}{4}$ 2.564(3) Å

(ii: xp1, y, zp1) (Fig. S9) and all the bridging Bi-O distances are longer with respect to the not-bridging Bi-O bonds (2.375 (4) Å av.). The main difference between the two dimeric molecules is the presence of two hydrogen atoms connected to two oxygens of the non-bridging pydc fragments in one of them i.e. Bi1, and consequently the corresponding C1AeO2A distance (1.253 (6) Å) results slightly elongated with respect to the analogue in the other pydc fragments (C]O 1.232 Å av.). The pydc anions exhibit (02110), (01111) and (01110) modes (Fig. S10) [67]. This dimeric compound lies about an inversion centre. The distances between the two bismuth atoms in each dimer are too large for considering a bonding interaction (Bi1eBi1A 4.299(1) Å, Bi2eBi2A 4.214(1) Å) and they agree to similar dimeric complexes (BiBi 4.320 Å) [16]. In the structure of **3**, the values of 2.84 and 2.98 for s_i obtained for

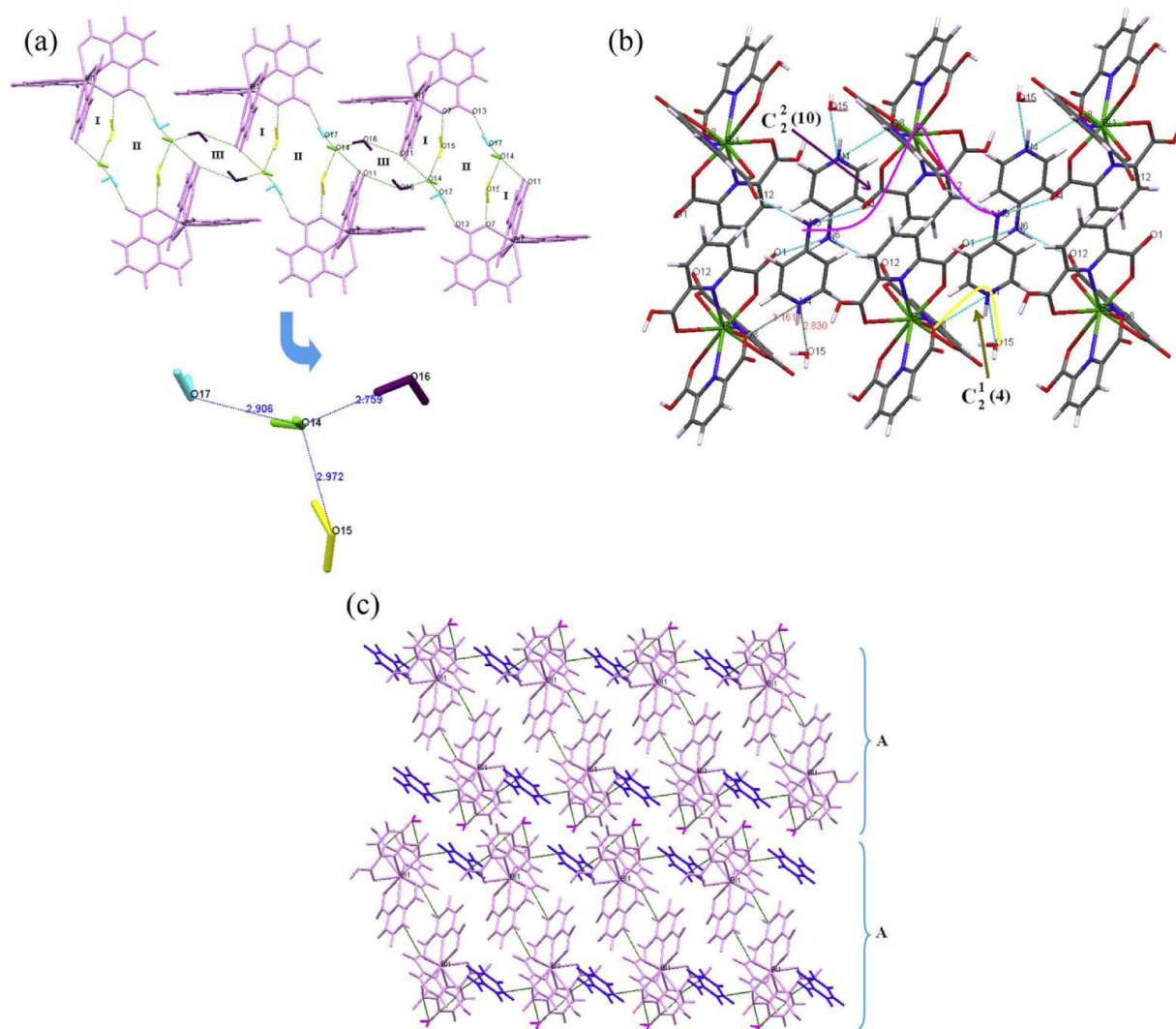


Fig. 5. (a) Graph set descriptors made by OeH/O interactions between water molecules and $[\text{Bi}(\text{pydc})_2(\text{pydcH})]_2$ with a view of the triangle shape tetrameric water cluster in 2, (b) A view of 1D chain constructed by NeH/O hydrogen bonds between anionic complexes, cationic fragments and uncoordinated water molecules in 2, (c) Hydrogen bonded layer A made by CeH/O interactions in 2.

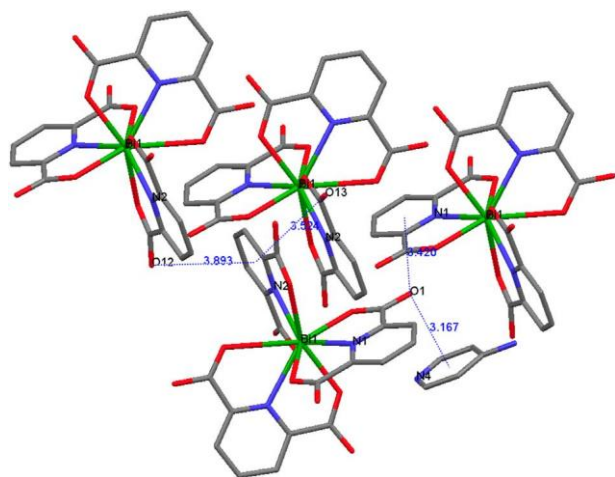


Fig. 6. CeO/p stackings interactions in compound 2, between C7eO1 and N4/ C25eC29 (x, y, z) ring, C7eO1 and N1/C1eC5 (1 x,1 y,1 z) ring, C23eO12 and N2/ C18eC22 (x,2 y,1 z) ring and C17eO13 and N2/C18eC22 (1 x,2 y,1 z) ring with the distances of 3.167(3), 3.420(3), 3.893(3) and 3.524(3)Å, respectively.

the Bi1 and Bi2 cations, respectively, are also in good agreement with the expected charge value of 3.00 (Table 4).

There are extensive intermolecular O-H...O, N-H...O and weak C-H...O hydrogen bonds, which cause the stability of the crystal structure (Table 3). The dimeric neutral and anionic molecules are disposed through alternate rows in the [010] direction, leading to the formation of an infinite three-dimensional H-bonded layer structure parallel to the (001) plane, in which all the molecules are directly connected by two strong (O21-H21A/O2B 2.685(6) Å and O11-H11B/O4D 2.689(6) Å) and two weaker (C5D-H5D/O4B 2.463 Å and C3B-H3B/O2D 2.581 Å) hydrogen bonds and also some π - π stacking interactions between the pyridine rings along the neutral molecules rows N1A/C2A-C6A rings with the distance of 3.903(3) Å (centroid-centroid distances) (Fig. 8a and b). The molecules of different planes are connected each other by the short hydrogen bond C4D-H4D/O2C (3.028(8) Å). Thus the molecules of the complex are directly interconnected in all directions of the crystal, without involving the pipzearH cation and the free water molecules. The volume between the (001) planes is filled by the pipzearH cations, connected to the dimeric molecules through

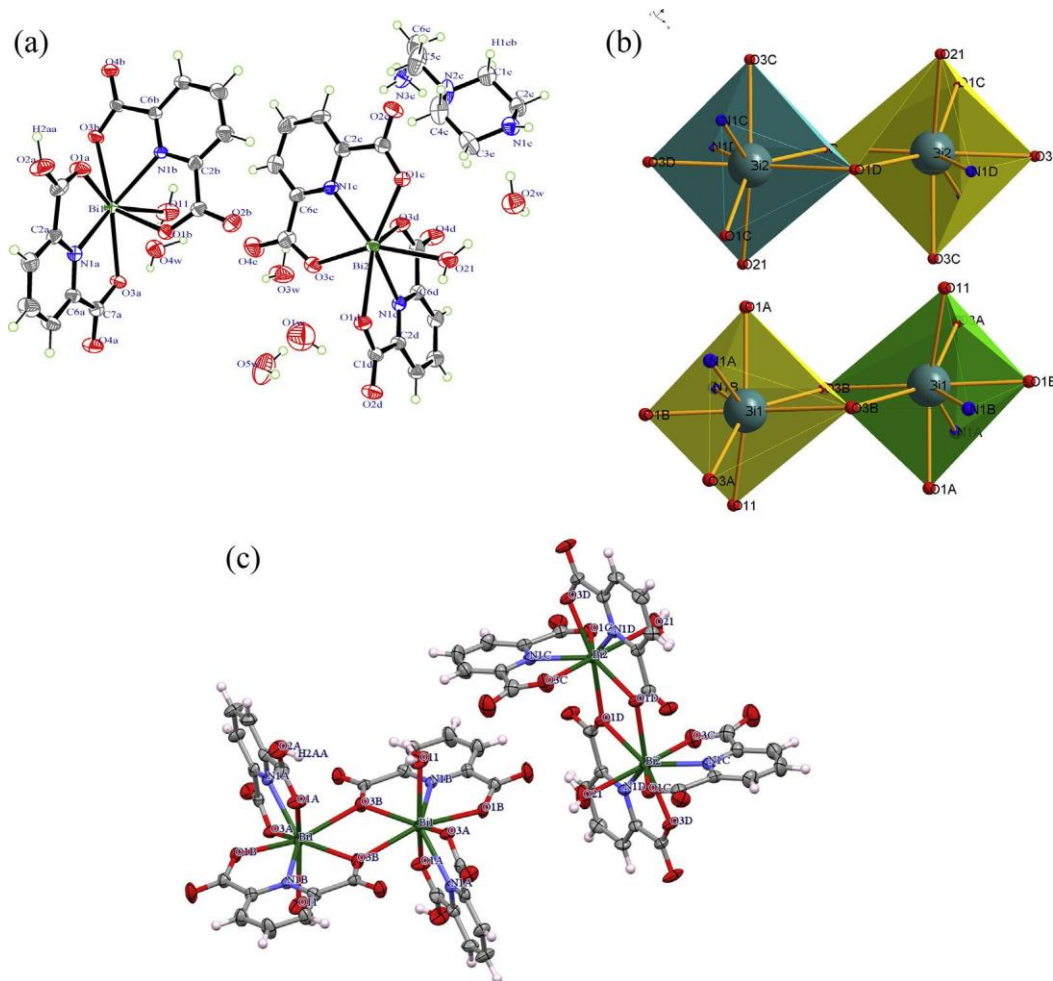


Fig. 7. (a) Asymmetric unit content of (pipzeaH)[Bi₂(pydc)₃(pydcH)(H₂O)₂].5H₂O, 3, showing the atom-labeling. Ellipsoids are at the 40% probability level; (b) Polyhedron showing the bicapped octahedral coordination environment, BiN₂O₆, at bismuth centers in 3; (c) Various coordination modes exhibited by pydc and pydcH anions in compound 3.

NeH/O, OeH\$\$\$N and CeH/O bonding.

Noteworthy is the strong hydrogen bond that involves the hydrogen connected to the oxygen of one pydc fragment and the positive charged N1E of the pipzeaH cation (O2AeH2AA/N1E 2.822(7) Å) (Fig. 8b). Clearly the acid H2AA is strongly attracted from the positive charge of the N1E atom, and this is the reason for the possibility to easily localize this hydrogen with X-ray diffraction; the consequence of this strong interaction is that the hydrogen connected to the N1E is very near to H2AA (H1EA/H2AA ¼ 1.61 Å) and thus in this case the hydrogen repulsion is negligible with respect to the electrostatic force connecting H2AA with N1E.

Finally, between the (001) planes are disposed also the free water molecules that contribute to the stabilization of the crystal packing, bonding both the pydc, the coordinated water ligands and the pipzeaH cations, through a great number of strong and weak hydrogen bonds (Table 3). As shown in Fig. 8c, the water molecules are involved in intermolecular OeH/O hydrogen bonding, and form a one-branched chain of the water molecules with the formula of (H₂O)₆.

4. Conclusion

Three new metal-organic compounds containing bismuth and pyridine-2,6-dicarboxylate (pydc), (2-apyH)₂[Bi(pydc)₂(pydcH)]\$ 2H₂O, 1, (4-apyH)[Bi(pydc)(pydcH)₂]\$4H₂O, 2 and (pipzeaH) [Bi₂(pydc)₃(pydcH)(H₂O)₂]\$5H₂O, 3 (pydc ¼ pyridine-2,6-dicarboxylate, 2-apy ¼ 2-aminopyridine, 4-apy ¼ 4-aminopyridine, pipzea ¼ 2-piperazin-1-ylethanamine), with zero-dimensional structures have been synthesized via proton transfer methodology and structurally characterized. The proton transfer methodology is utilized as a useful method for the preparation of anionic complexes. Variation in the coordination and H-bonded structure is achieved by varying the amines in the synthesis. Compounds 1 and 2 display similar bismuth coordination environments with a stereochemically inactive lone pair of bismuth. Also four different coordination modes of metal-dipicolinate exist in 1, 2 and 3. Comparison of our results in the solid state confirms the theory that complexation behavior of bismuth dipicolinate in the same synthetic conditions depends on using different organic amines as counter cation precursors leading to different structural entities. Furthermore, four different coordination modes of metal-dipicolinate exist in 1, 2 and 3, of which two are rare coordination types; one involving metal coordination to the short C]O(BiO) bond of the protonated acid, which is quite rare in the bismuth

dipicolinate complexes; the other which involves metal coordination to a protonated oxygen atom; that has not yet been reported in the bismuth dipicolinate complexes. Fascinating crystal packing

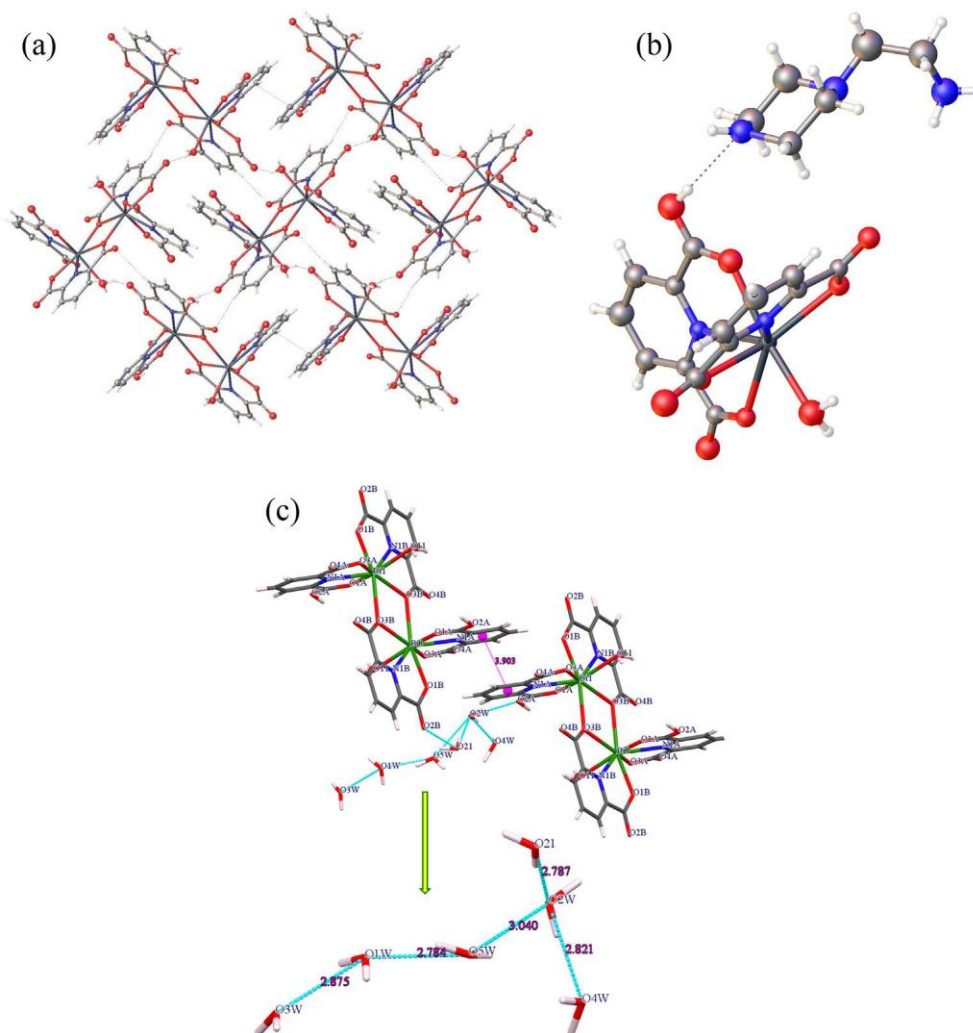


Fig. 8. (a) View of the (001) plane of the crystal packing of **3**. Hydrogen bonds and aromatic stacking π - π interactions are shown with dashed lines; (b) the strong hydrogen bond that involves the hydrogen connected to the oxygen of one pydc fragment and the positive charged N1E of the pipzeaH cation in **3**; (c) π - π stacking interactions and a view of the hexameric water cluster in **3**, the π - π stacking between N1a/C2a-C6a (x, y, 1 z) rings with the distance of 3.903(3) Å.

motif results from the concert of intermolecular interactions, such as electrostatic attraction between ion pairs, different kinds of hydrogen bonding (OeH/O, NeH/O, CeH/O) and stacking interactions.

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Supplementary data

Crystallographic data for the three structures have been deposited with the Cambridge Crystallographic Data Centre, CCDC 874050 for compound **1**, CCDC 874056 for compound **2** and CCDC 918173 for compound **3**. Copies of the data can be obtained free of charge on application to the Director, CCDC, 12 Union Road, Cam-bridge CB2 1EZ, UK (Fax. þ44 1223 336033; e-mail: deposit@ccdc.cam.ac.uk or www: <http://www.ccdc.cam.ac.uk>). Plots of NMR,

UV and TGA spectra together with additional figures are also available in supplementary data.

Appendix A. Supplementary data

Supplementary data related to this article can be found at [http:// dx.doi.org/10.1016/j.molstruc.2015.06.084](http://dx.doi.org/10.1016/j.molstruc.2015.06.084).

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