

Molecular Engineering to Tune the Ligand Environment of Atomically Dispersed Nickel for Efficient Alcohol Electrochemical Oxidation

Zhifu Liang, Daochuan Jiang, Xiang Wang, Mohsen Shakouri, Ting Zhang, Zhongjun Li, Pengyi Tang,* Jordi Llorca, Lijia Liu,* Yupeng Yuan, Marc Heggen, Rafal E. Dunin-Borkowski, Joan R. Morante, Andreu Cabot,* and Jordi Arbiol*

Atomically dispersed metals maximize the number of catalytic sites and enhance their activity. However, their challenging synthesis and characterization strongly complicates their optimization. Here, the aim is to demonstrate that tuning the electronic environment of atomically dispersed metal catalysts through the modification of their edge coordination is an effective strategy to maximize their performance. This article focuses on optimizing nickel-based electrocatalysts toward alcohol electrooxidation in alkaline solution. A new organic framework with atomically dispersed nickel is first developed. The coordination environment of nickel within this framework is modified through the addition of carbonyl (C=O) groups. The authors then demonstrate that such nickel-based organic frameworks, combined with carbon nanotubes, exhibit outstanding catalytic activity and durability toward the oxidation of methanol (CH₃OH), ethanol (CH₃CH₂OH), and benzyl alcohol (C₆H₅CH₂OH); the smaller molecule exhibits higher catalytic performance. These outstanding electrocatalytic activities for alcohol electrooxidation are attributed to the presence of the carbonyl group in the ligand chemical environment, which enhances the adsorption for alcohol, as revealed by density functional theory calculations. The work not only introduces a new atomically dispersed Ni-based catalyst, but also demonstrates a new strategy for designing and engineering high-performance catalysts through the tuning of their chemical environment.

1. Introduction

The electrochemical oxidation of alcohols is at the basis of several strategic chemical processes and energy technologies, including electrochemical organic synthesis, biomass, and organic waste valorization, and direct alcohol fuel cells (DAFC).^[1] The electrooxidation of alcohols offer numerous advantages over competing valorization strategies and energy technologies, such as high energy densities and energy conversion efficiency, operation/reaction at low working temperature with high selectivity, and low fuel cost with easy production, transportation, and storage.^[1] However, despite their numerous advantages, the deployment of electrocatalytic valorization strategies and DAFCs is hampered by their low cost-effectiveness in large part associated with the use of moderately efficient catalysts based on high-cost and low-abundance elements such as platinum and palladium.^[2,3] Besides, current alcohol electrooxidation catalysts become rapidly poisoned with carbon monoxide,

Z. F. Liang, T. Zhang, J. Arbiol
Catalan Institute of Nanoscience and Nanotechnology (ICN2)
CSIC and BIST
Campus UAB, Bellaterra, Barcelona, Catalonia 08193, Spain
E-mail: arbiol@icrea.cat

Z. F. Liang, X. Wang, A. Cabot, J.R. Morante
Functional Nanomaterials Group
Catalonia Institute for Energy Research – IREC
Sant Adrià de Besòs, Barcelona Catalonia 08930, Spain
E-mail: acabot@irec.cat

 The ORCID identification number(s) for the author(s) of this article can be found under <https://doi.org/10.1002/adfm.202106349>.

© 2021 The Authors. Advanced Functional Materials published by Wiley-VCH GmbH. This is an open access article under the terms of the Creative Commons Attribution-NonCommercial-NoDerivs License, which permits use and distribution in any medium, provided the original work is properly cited, the use is non-commercial and no modifications or adaptations are made.

DOI: 10.1002/adfm.202106349

D. Jiang, Y. Yuan
School of Chemistry and Chemical Engineering
Anhui University
Heifei Anhui 230601, China

M. Shakouri
Canadian Light Source
Saskatoon S7N 0X4, Canada

Z. J. Li
School of Physics
Hefei University of Technology
Hefei 230009, P. R. China

P. Y. Tang
State Key Laboratory of Information Functional Materials, 2020 X-Lab
Shanghai Institute of Microsystem and Information Technology
Chinese Academy of Sciences
Shanghai 200050, China
E-mail: py.tang@mail.sim.ac.cn

which further limits their cost-effectiveness. Thus, the development of efficient and stable alcohol oxidation electrocatalysts based on abundant elements is a topic of major interest.

Several non-precious transition metal catalysts have been proposed for alcohol electro-oxidation.^[4,5] Among them, nickel-based catalysts are considered the most promising candidates owing to their high catalytic activity, low price, and notable abundance. Numerous Ni-based catalysts, such as Ni_{0.75}Cu_{0.25},^[6] NiSn alloy,^[7] Ni₉₃Bi₇,^[8] NiSe,^[9] NiO,^[10] and metal–organic frameworks with Ni(OH)₄,^[11] and Ni–Co hydroxide^[12] have been developed and tested for alcohol electrochemical-oxidation. However, the reported onset potentials and stabilities are still not fully satisfactory, which calls for novel strategies to design and engineer nickel-based catalysts with improved performance.

Recently, 2D covalent organic frameworks and metal–organic frameworks (MOF) with highly tunable pore structure and functionality, and high density of active sites, have arisen broad interest as electrocatalysts.^[13] Optimized electrocatalysts based on such 2D structures have demonstrated outstanding performances toward oxygen evolution and reduction,^[14,15] and carbon dioxide reduction.^[16,17] However, despite their high potential for alcohol electrooxidation, these 2D materials and particularly a nickel-containing 2D organic framework have yet to be developed and optimized toward this electrocatalytic reaction.

Here, we report the synthesis of an atomically-dispersed nickel 2D molecular organic framework containing abundant carbonyl (C=O) functional groups within the nickel ion coordination environment. Its atomic structure and how the outstanding electrocatalytic activities and durabilities obtained for the electrochemical-oxidation of methanol, ethanol, and benzyl alcohol were analyzed, are related to the alcohol adsorption ability of the carbonyl functional groups and the high unsaturated coordination number of the nickel ions within the new catalyst presented here.

2. Results and Discussion

The nickel-based 2D organic framework with abundant carbonyl functional groups (Ni–2D–O–SA) was solvothermally synthesized through a Schiff-based reaction between

2,5-hydroxyterephthalaldehyde (HBC) (Scheme S1, Supporting Information) and tetramino-benzoquinone (TABQ) (Scheme S2, Supporting Information) in *N*-methyl-2-pyrrolidone (NMP) (see details in **Figure 1a**; Scheme S3, Supporting Information). The dark-black color of the obtained solid indicated the formation of a conjugated polymer, which was washed by water and methanol to remove low molecular mass impurities, and then dried under vacuum. For comparison, a reference nickel-based 2D organic framework without carbonyl chemical groups (Ni–2D–SA) was prepared under the same reaction condition but using 1,2,4,5-benzenetetraamine tetrahydrochloride (TAB) instead of TABQ (Scheme S4, Supporting Information).^[18]

Scanning electron microscopy (SEM) images (**Figure S1a,b**, Supporting Information) show that the Ni–2D–O–SA sample has a cotton-like morphology. The X-ray diffraction (XRD) pattern (**Figure S3**, Supporting Information) of the Ni–2D–O–SA powder was similar to that of graphite, with a diffraction peak at 26.29° corresponding to the (001) family planes and which indicated π – π layered stacking.^[19] In contrast to the clear crystalline and layered structure displayed by Ni–2D–O–SA, the XRD pattern of Ni–D–SA displayed a much broader and less intense peak at around 26°. This result indicates that the introduction of carbonyl groups enhances the crystallinity of the nickel-based 2D organic framework used here.

Fourier transform infrared (FT-IR) spectroscopy was used to evaluate the chemical structure of the samples (**Figure S4**, Supporting Information). Based on the FT-IR spectra, the signal for C=N appeared at 1642 cm⁻¹ for Ni–2D–SA, and 1648 cm⁻¹ for Ni–2D–O–SA. The vibration of the carbonyl bond (C=O) in Ni–2D–O–SA negatively shifted below 1600 cm⁻¹ and overlapped with other vibrations. For Ni–2D–O–SA, the vibration peak for the C–(C=O)–C bond, which is a footprint of TABQ, shifted from 1140 to 1010 cm⁻¹ due to the coordination effect and the attraction between the layers. This result already points out to the successful introduction of the carbonyl groups in the organic framework.^[20]

The survey X-ray photoelectron spectroscopy (XPS) spectra revealed the sole presence of C, N, Ni, and O in both samples (**Figures S5 and S6**, Supporting Information).^[21] For both samples, the high-resolution Ni 2p spectrum showed a unique doublet at binding energies of 856 eV (Ni 2p_{3/2}) and 873.3 eV (Ni 2p_{1/2}), significantly above those of NiO.^[22,23]

The local environment in Ni–2D–O–SA was further analyzed and compared with that of NiO using X-ray absorption fine structure (XAFS) at the Ni K-edge (**Figure 1b–d**; **Figure S7**, Supporting Information). Comparing the spectral features at the near-edge, that is, the X-ray absorption near-edge structure (XANES), we observe the absorption onset of Ni–2D–O–SA at slightly lower energy than that of NiO (**Figure 1b**). Differences are more clearly seen when plotting the first derivative of the two spectras (**Figure S7a**, Supporting Information), as shown in the inset of **Figure 2b**. Besides, the main absorption peak of Ni–2D–O–SA has a lower intensity than that of NiO. Both features indicate that compared with NiO, where Ni is fully surrounded by O, within Ni–2D–O–SA, Ni atoms are partially surrounded by a lower electronegativity element, that is, N. A more quantitative picture was obtained by analyzing the extended XAFS (EXAFS) features. The fitting of the Fourier transformed EXAFS spectrum suggests that Ni is coordinated with two N atoms and two O atoms, at a bond length 1.87 and 2.05 Å,

P. Y. Tang, M. Heggen, R. E. Dunin-Borkowski
Ernst Ruska-Centre for Microscopy and Spectroscopy with Electrons and Peter Grünberg Institute, Forschungszentrum Jülich GmbH
52425 Jülich, Germany

J. Llorca
Institute of Energy Technologies
Department of Chemical Engineering and Barcelona
Research Center in Multiscale Science and Engineering
Universitat Politècnica de Catalunya
EEBE
Barcelona Catalonia 08019, Spain

L. Liu
Department of Chemistry
Western University
1151 Richmond Street, London ON N6A5B7, Canada
E-mail: lijia.liu@uwo.ca

A. Cabot, J. Arbiol
ICREA
Pg. Lluís Companys 23, Barcelona Catalonia 08010, Spain

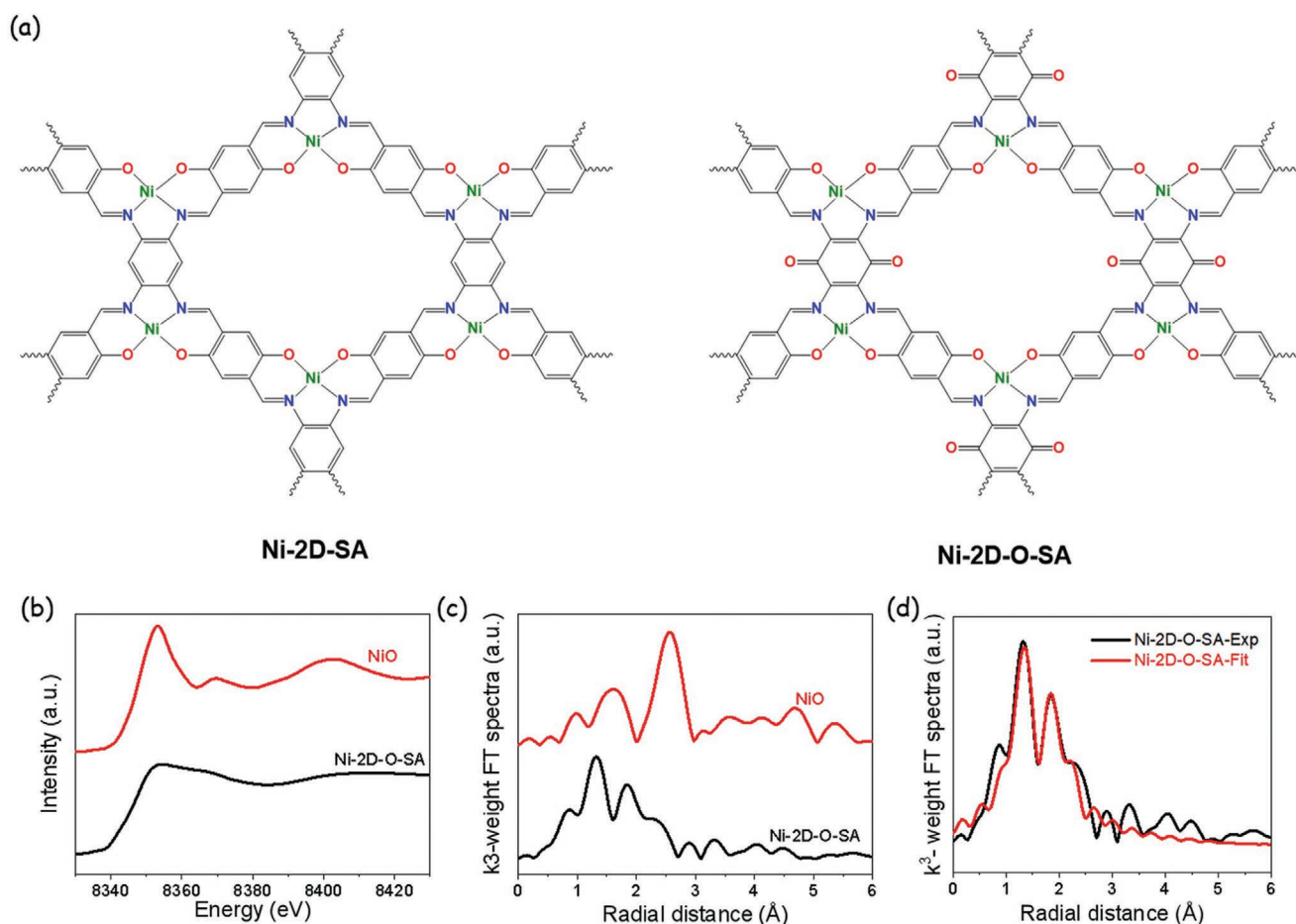


Figure 1. a) Scheme of the synthesis procedure used to produce Ni-2D-O-SA. b) Ni K-edge XANES spectrum of Ni-2D-O-SA compared with a commercial NiO powder. c) Fourier transformed Ni K-edge EXAFS spectra of Ni-2D-O-SA and NiO plotted in R-space. d) Fitting of the Fourier transformed EXAFS spectrum in R-space of Ni-2D-O-SA.

respectively (Figure 1c,d).^[24] The contribution at longer radial distances comes from C atoms with a coordination number of 6 and a bond length of 2.65 Å. The detailed fitting parameters can be found in Table S1, Supporting Information. Overall, XAFS analysis provided strong evidence of the formation of the NiN₂O₂ salophen structure unit in the Ni-2D-O-SA framework.

High angle annular dark-field (HAADF) aberration-corrected scanning transmission electron microscopy (STEM) was used to analyse the nickel dispersion. As shown in Figure 2a–c; Figure S8a–c, Supporting Information, Ni-2D-O-SA contained a high density of isolated and homogeneously distributed nickel atoms. The uniform distribution of Ni, C, N, and O was further confirmed by energy-dispersed X-ray spectroscopy (EDS) elemental mapping (Figure 2d).

The surface area of Ni-2D-O-SA and Ni-2D-SA, evaluated from N₂ adsorption–desorption isotherms at 77 K (Figure S9, Supporting Information) using the Brunauer–Emmett–Teller equation was 21.1 and 83 m² g^{−1}, respectively. Thus, the introduction of carbonyl functional groups significantly decreased the surface area of the material, which we associate with an enhancement of the layer–layer interaction that results in slightly more densely packed frameworks.

To improve the electrical conductivity of these nickel-based organic frameworks, as required to apply them as

electrocatalysts, they were loaded on multi-wall carbon nanotubes (CNT) through the π – π packing interaction. The XRD patterns of the obtained Ni-2D-O-SA–CNT and Ni-2D-SA–CNT composites, showed a unique peak at $\approx 26^\circ$ (Figure S10, Supporting Information), which was assigned to the overlap between the CNT and the organic framework structures. SEM (Figure S1c,d, Supporting Information), STEM images, and EDX compositional maps (Figure S11, Supporting Information) showed the composites morphology and confirmed the homogeneous loading of Ni-2D-O-SA on the CNTs.

Figure 3a exhibits the cyclic voltammograms (CV) of the two samples in the range from 0 to 0.65 V in 1.0 M KOH electrolyte with a scan rate of 50 mV s^{−1}. The anodic peak at 0.50 V versus Hg/HgO measured in the forward scan with Ni-2D-O-SA–CNT is attributed to Ni²⁺ oxidation, and the cathodic peak in the backward scan at 0.38 V corresponds to the reduction of Ni³⁺ species. For Ni-2D-SA–CNT, the anodic peak shifted to 0.54 V and the cathodic peak to 0.42 V. The current density at the Ni²⁺ oxidation peak in the Ni-2D-O-SA–CNT electrocatalyst was 19 mA cm^{−2}, which is ≈ 9.5 times higher than that of Ni-2D-SA–CNT, 2 mA cm^{−2}. These results indicate a much higher electrochemical activity of the Ni ions within Ni-2D-O-SA–CNT and their easier oxidation to Ni³⁺. Besides, the ratio between the anodic and the cathodic peak current

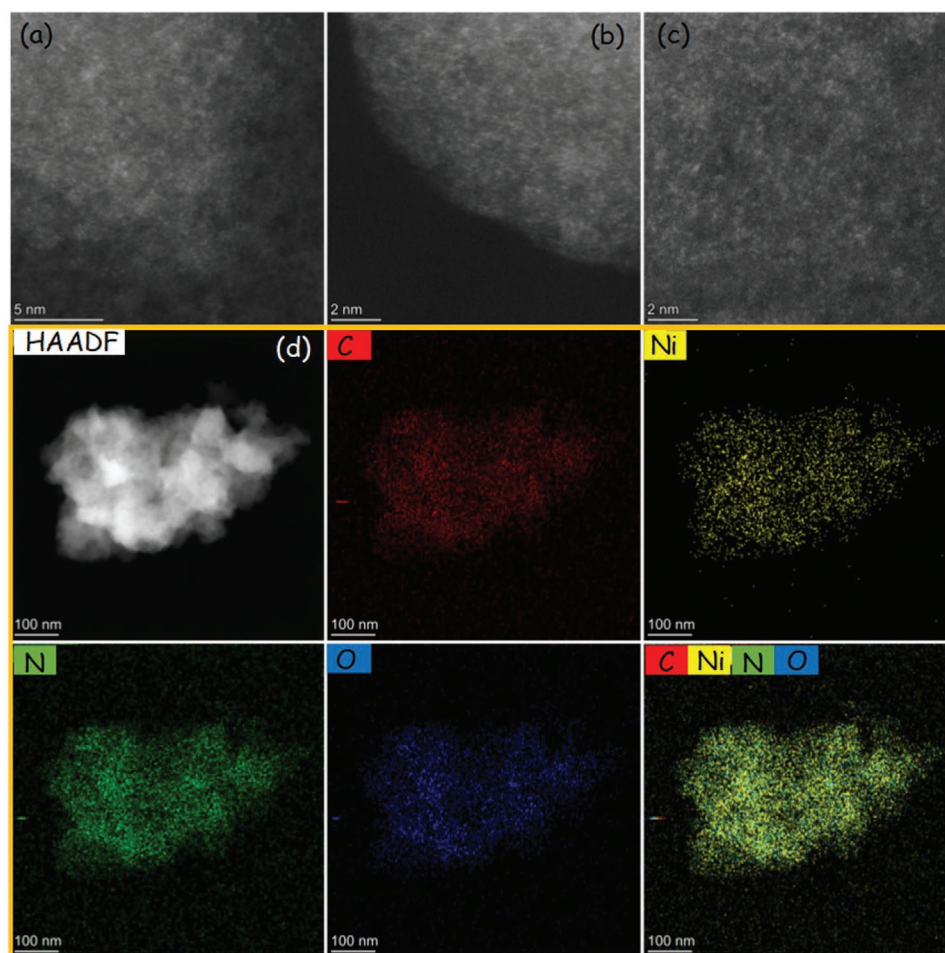


Figure 2. a–c) HAADF–STEM images of Ni–2D–O–SA displaying the presence of atomically dispersed nickel atoms. d) Low magnification HAADF–STEM image and EDS elemental mapping.

densities was close to 1 for both catalysts, indicating a reversible $\text{Ni}^{2+}/\text{Ni}^{3+}$ redox reaction in both materials.^[25]

CV of the reduction/oxidation process allowed to estimate the surface coverage (Γ) of $\text{Ni}^{2+}/\text{Ni}^{3+}$ within Ni–2D–O–SA–CNT to be $1.32 \times 10^{-7} \text{ mol cm}^{-2}$, two orders of magnitude above that of Ni–2D–SA–CNT, $5.17 \times 10^{-9} \text{ mol cm}^{-2}$ (Figure 3a; Equation (S1), Supporting information).^[26] Figure 3b; Figure S12, Supporting Information, display CV curves at different scan rates, from 10 to 100 mV s^{-1} , obtained with these two catalysts in 1.0 M KOH electrolyte. The variation of the anodic and cathodic peak current densities with the square root of the voltage scan rate was linearly fitted with $R^2 > 0.99$ for both samples (Figure S12, Supporting Information), demonstrating the $\text{Ni}^{2+}/\text{Ni}^{3+}$ redox reaction to be diffusion limited.^[25] From this dependence, an apparent diffusion coefficient D' was evaluated (see Equation S3, Supporting Information). The D' obtained with Ni–2D–O–SA–CNT is $2.3 \times 10^{-6} \text{ g cm}^{-1} \text{ s}^{-1}$, which is close to an order of magnitude higher than with Ni–2D–SA–CNT, $3.3 \times 10^{-7} \text{ g cm}^{-1} \text{ s}^{-1}$.

The alcohol electrooxidation performance of Ni–2D–O–SA–CNT and Ni–2D–SA–CNT electrocatalysts was analyzed by linear sweep voltammetry (LSV) in 1 M KOH and 1 M KOH + 1 M methanol solutions using a conventional three-electrode setup (Figure 3c,d). The OER activity of Ni–2D–O–SA–CNT was just

slightly improved compared to that of Ni–2D–SA–CNT. However, the Ni–2D–O–SA–CNT catalysts showed much higher activity toward methanol oxidation reaction (MOR), with current densities increasing from 31 to 106 mA cm^{-2} in the range 0.5–0.6 V versus Hg/HgO, that is, six-fold higher current densities compared with Ni–2D–SA–CNT. Thus, the introduction of the carbonyl group (C=O) at the ligand environment of the dispersed Ni-atoms active sites clearly results in much higher MOR catalytic activities.

The charge transfer resistances of these two catalysts were evaluated through electrochemical impedance spectroscopy (EIS). The EIS Nyquist plot of Ni–2D–O–SA–CNT (Figure 3e) presents a smaller semicircle than Ni–2D–SA–CNT at 0.56 V in 1 M KOH + 1 M methanol electrolyte, denoting a lower charge transfer resistance (R_3) and thus faster MOR kinetics.^[27] On the other hand, the charge transfer resistance for OER of Ni–2D–O–SA–CNT is similar to that of Ni–2D–SA–CNT at 0.67 V (Figure S13, Supporting Information), which is consistent with the similar OER activities of these two catalysts.

As shown in the chronopotentiometry (CP) curve (Figure 3f), the Ni–2D–O–SA–CNT electrocatalyst maintained constant current densities of 10 and 30 mA cm^{-2} for 12 h without a

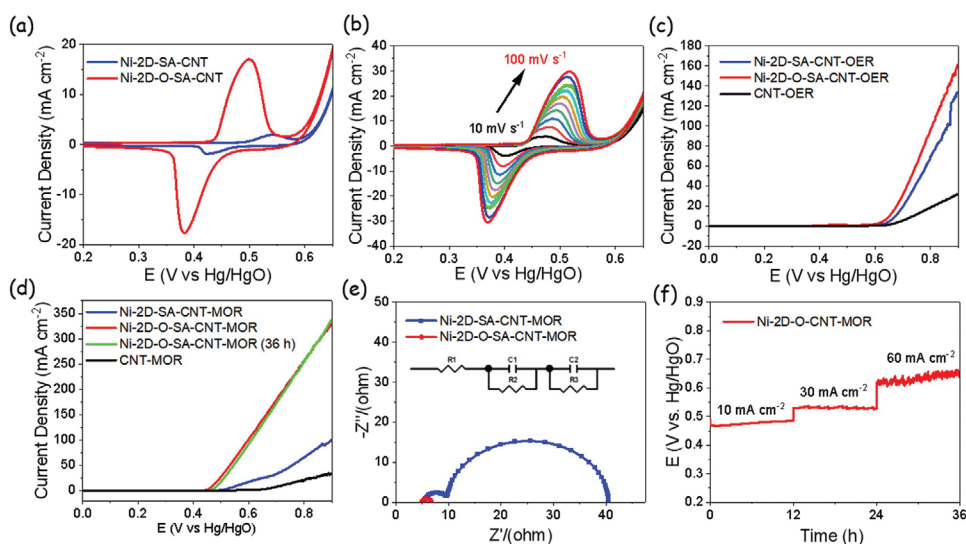


Figure 3. a) CV of Ni-2D-O-SA-CNT and Ni-2D-SA-CNT electrocatalysts in 1 M KOH at 50 mV s⁻¹ potential sweep rate. b) CVs of Ni-2D-O-SA-CNT electrocatalysts in 1 M KOH obtained at an increasing potential sweep rate, from 10 to 100 mV s⁻¹. c,d) LSV curves obtained with Ni-2D-SA-CNT, Ni-2D-O-SA-CNT, and reference carbon nanotubes (CNTs) catalysts in 1.0 M KOH (c) and 1 M KOH + 1 M methanol in the potential range 0 to 0.9 V versus Hg/HgO at a scan rate of 5 mV s⁻¹ (d). e) EIS Nyquist plots at 0.56 V in 1 M KOH + 1 M methanol solution. f) CP long-term stability of Ni-2D-O-SA-CNT at 10, 30, and 60 mA cm⁻². No iR compensation was applied to the displayed data.

significant increase in the required potential. When increasing the current density to 60 mA cm⁻², the potential raised ≈4.8% during 12 h, which is attributed to the methanol consumption. After 36 h chronopotentiometry test, LSV curves recorded with the used electrocatalysts in a fresh 1 M KOH + 1 M methanol electrolyte (Figure 3d) showed almost no variation with respect to those obtained with the fresh electrocatalysts, what demonstrates the excellent stability of Ni-2D-O-SA-CNT during MOR. After the stability test, the pH value of the solution is 14.12, which is not a significant change compared to the initial solution (pH = 14).

The XRD pattern of Ni-2D-O-SA-CNT after MOR (Figure S14, Supporting Information) displayed no significant variation with respect to the sample before the electrochemical tests, which confirmed the excellent structural stability of the 2D material and the growth of no additional crystalline phase of the nickel or nickel oxide nanoparticles during the reaction, indicating there is no agglomeration of nickel atom active sites.

Methanol adsorption is the initial and a limiting step in MOR. To understand the contribution of C=O groups to methanol adsorption, density functional theory (DFT) calculations were carried out. Selected fragments were shown in Figure S15, Supporting Information. Figure 4 shows the optimal methanol adsorption configurations on top of a Ni site in both Ni-2D-SA and Ni-2D-O-SA. The adsorption energy is -0.248 eV and -0.225 eV in Ni-2D-SA and Ni-2D-O-SA, respectively. The negative adsorption energies suggest that methanol adsorption is thermodynamically favored in both cases. Interestingly, the adsorption distance (*d*) between the O atom of methanol and Ni in Ni-2D-O-SA is 3.11 Å, which is shorter than in Ni-2D-SA (3.27 Å). Compared with Ni-2D-SA, the 10% lower adsorption energy and 5% shorter adsorption distance indicate a more effective adsorption of the methanol molecule on the surface of

Ni-2D-O-SA. What is more, upon adsorption, the H-O-CH₃ angle in Ni-2D-O-SA increases from 108.25° to 109.04°, while the H-O-CH₃ angle in Ni-2D-SA remains almost unchanged after optimization. This result indicates that Ni-2D-O-SA can facilitate the activation of the methanol molecule. We also identify the O of the C=O group as a methanol adsorption site. DFT calculations show that the methanol molecule tends to adsorb at such O site with the OH group in methanol parallel to the surface. The shortest distance between the methanol molecule and the surface is 1.87 Å, which is much smaller than in the Ni-2D-SA surface (2.5 Å). Thus, based on the DFT calculations, it is reasonable to conclude that the addition of carbonyl (C=O) groups to the organic framework results in more effective adsorption as well as activation of the methanol molecule, which explains the higher activity of Ni-2D-O-SA toward MOR when compared with Ni-2D-SA.

Beyond MOR, we also examined the electrocatalytic activity of the produced catalysts toward the oxidation of other alcohols. Compared with Ni-2D-SA-CNT, Ni-2D-O-SA-CNT presented significantly higher electrocatalytic performance toward the electrochemical oxidation of both ethanol (EOR) (Figure 5a) and benzyl alcohol (BOR) (Figure 5b). Especially for the BOR, the current density at 0.6 V obtained with Ni-2D-O-SA-CNT was ≈77 mA cm⁻², which is ≈eight times higher than that of Ni-2D-SA-CNT (9.7 mA cm⁻²).

CP measurements were used to test the stability of Ni-2D-O-SA-CNT for EOR. As shown in Figure S16, Supporting Information, Ni-2D-SA-CNT exhibited excellent long-term stability for EOR, similar to that obtained for MOR. For BOR, chronoamperometric measurements (Figure S17, Supporting Information) show the current density decay associated with the consumption of benzyl alcohol. Notice in this respect that the initial concentration of benzyl alcohol was just 0.1 M, since the steric hindrance of benzene may prevent an efficient

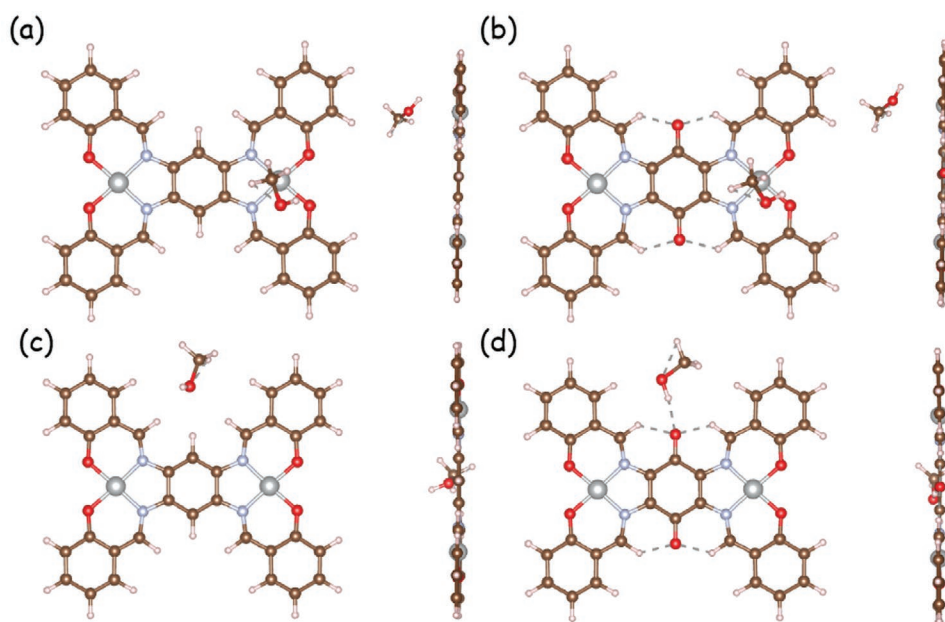


Figure 4. a,b) Optimal methanol adsorption configurations on top of the Ni site in both Ni-2D-SA (a) and Ni-2D-O-SA (b). c,d) Optimal methanol adsorption configurations on top of the carbonyl site in both Ni-2D-SA (c) and Ni-2D-O-SA (d) (red is oxygen, light blue is nitrogen, grey is nickel, brown is carbon).

reaction at higher concentrations. The EIS spectra for EOR and BOR were similar to those obtained for MOR, confirming that Ni-2D-O-SA-CNT is also characterized by a highly efficient charge transfer and a highly favorable EOR and BOR kinetics (Figure S18, Supporting Information).

Finally, in order to confirm the product of the electrochemical oxidation of alcohol, nuclear magnetic resonance (NMR) tests were conducted to analyze the associated products of electrochemical oxidation of alcohol. As shown in Figure S19, Supporting Information, the NMR spectrum after the MOR stability test displays the characteristic chemical shift of HCOO^- at 8.33 ppm, while the characteristic chemical shift of HCHO in the range of $\approx 9\text{--}10$ ppm was not observed, demonstrating that the final product of methanol oxidation is formate (HCOO^-). Furthermore, the NMR spectrum

(Figure S20, Supporting Information) of the products from benzyl alcohol oxidation showed a chemical shift in the range of $\approx 7\text{--}8$ ppm, which is assigned to the characteristic chemical shift of the proton in the benzene ring of benzoate. The characteristic chemical shift belonging to the aldehyde ($-\text{CHO}$) of benzaldehyde was not observed at the range of $\approx 9\text{--}10$ ppm, demonstrating that the final product of the benzyl alcohol oxidation was benzoate ($\text{C}_6\text{H}_5\text{COO}^-$). The NMR spectra of the EOR electrolyte did also not display the characteristic chemical shift of the aldehyde ($-\text{CHO}$). The characteristic chemical shift of CH_3COO^- appears at 1.83 ppm (Figure S21, Supporting Information).

Taking into account the above NMR results, we propose the following electrochemical reaction for the alcohol electrochemical oxidation (Figure 6):

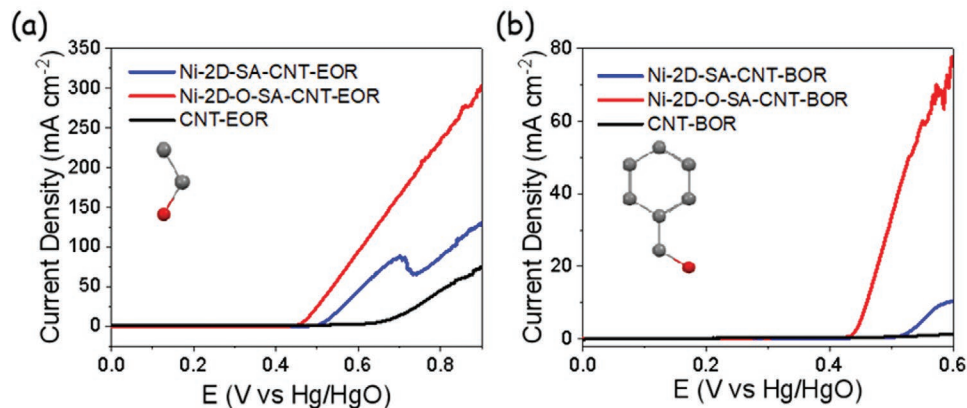


Figure 5. a) LSV of Ni-2D-SA-CNT, Ni-2D-O-SA-CNT, and CNT in 1 M KOH+1 M methanol from 0 to 0.9 V versus Hg/HgO at a scan rate of 5 mV s^{-1} . b) LSV of Ni-2D-SA-CNT, Ni-2D-O-SA-CNT, and the CNT in 1 M KOH + 0.1 M benzyl alcohol from 0 to 0.9 V versus Hg/HgO at a scan rate of 5 mV s^{-1} .

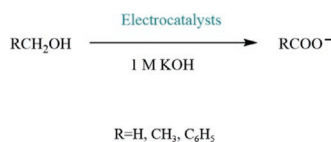


Figure 6. Scheme of the electrochemical oxidation reaction of alcohol.

3. Conclusion

In conclusion, we detailed the synthesis of a new nickel-based 2D organic framework and proposed a facile molecular engineering strategy to tune the ligand coordination environments of nickel ion centers. Through this strategy, carbonyl (C=O) groups were introduced and proved to decisively contribute to an efficient alcohol electro-oxidation in alkaline solution. EXAFS and XANES analysis confirmed the Ni salophen unit structure, NiN₂O₂, within the 2D organic frameworks. The atomically dispersed nickel within a framework containing the carbonyl groups displayed outstanding MOR, EOR, and BOR performance. DFT calculations suggested that the introduction of the C=O functional groups favor the absorption and activation of methanol molecules, which is at the origin of the improved MOR observed. The present work not only offers novel materials and a promising strategy to design more efficient atomically dispersed nickel-based catalysts for alcohol oxidation, but it could also benefit other research fields related to electrochemical-organic synthesis, energy conversion and storage.

Supporting Information

Supporting Information is available from the Wiley Online Library or from the author.

Acknowledgements

Z.F.L. and D.J. contributed equally to this work. ICN2 is supported by the Severo Ochoa program from Spanish MINECO (Grant No. SEV-2017-0706) and is funded by the CERCA Programme /Generalitat de Catalunya. Part of the present work has been performed in the framework of Universitat Autònoma de Barcelona Materials Science Ph.D. program. Z.L. acknowledges funding from MINECO SO-FPT Ph.D. grant (SEV-2013-0295-17-1). This project has received funding from the European Union's Horizon 2020 research and innovation program under grant agreement No 823717-ESTEEM3. The present work is supported by the European Regional Development Fund and by the Spanish MINECO through the projects ENE2016-77798-C4-3-R, ENE2017-85087-C3, and project NANOGEN (PID2020-116093RB-C43). X.W. and T.Z. thank the China Scholarship Council for the scholarship support. P. Tang acknowledges the Humboldt Research Fellowship. Authors acknowledge funding from Generalitat de Catalunya 2017SGR327 and 2017SGR1246. L.L. acknowledges the support from Natural Sciences and Engineering Research Council of Canada (NSERC, DG RGPIN-2020-06675). J.L. is a Serra Hünter Fellow and is grateful to MICINN/FEDER RTI2018-093996-B-C31, GC 2017 SGR 128, and to ICREA Academia program. XAS measurements were performed at the Canadian Light Source, a national research facility at the University of Saskatchewan, which is supported by the Canada Foundation for Innovation (CFI), the Natural Sciences and Engineering Research Council (NSERC), the National Research Council (NRC), the Canadian Institutes of Health Research (CIHR), the Government of Saskatchewan, and the University of Saskatchewan.

Conflict of interest

The authors declare no conflict of interest.

Data Availability Statement

Research data are not shared.

Keywords

alcohol oxidation, atomically dispersed metals, electrocatalytic oxidation, nickel, 2D organic frameworks

Received: July 2, 2021

Revised: August 12, 2021

Published online: September 16, 2021

- [1] a) E. Antolini, E. R. Gonzalez, *J. Power Sources* **2010**, *195*, 3431; b) K. C. Cheung, W. L. Wong, D. L. Ma, T. S. Lai, K. Y. Wong, *Coord. Chem. Rev.* **2007**, *251*, 2367; c) D. P. Hickey, M. S. McCammant, F. Giroud, M. S. Sigman, S. D. Minter, *J. Am. Chem. Soc.* **2014**, *136*, 15917; d) N. Kakati, J. Maiti, S. H. Lee, B. Viswanathan, Y. S. Yoon, *Chem. Rev.* **2014**, *114*, 12397; e) R. Ciriminna, G. Palmisano, M. Pagliaro, *ChemCatChem* **2015**, *7*, 552.
- [2] Y. Tong, X. Yan, J. Liang, S. X. Dou, *Small* **2021**, *17*, 1904126.
- [3] M. A. F. Akhairi, S. K. Kamarudin, *Int. J. Hydrogen Energy* **2016**, *41*, 4214.
- [4] a) S. Rezaee, S. Shahrokhian, *Appl. Catal., B* **2019**, *244*, 802; b) A. Badalyan, S. S. Stahl, *Nature* **2016**, *535*, 406.
- [5] G. M. K. Tolba, N. A. M. Barakat, A. M. Bastaweesy, E. A. Ashour, W. Abdelmoez, M. H. El-Newehy, S. S. Al-Deyab, H. Y. Kim, *J. Mater. Sci. Technol.* **2015**, *31*, 97.
- [6] X. Cui, P. Xiao, J. Wang, M. Zhou, W. Guo, Y. Yang, Y. He, Z. Wang, Y. Yang, Y. Zhang, Z. Lin, *Angew. Chem., Int. Ed.* **2017**, *56*, 4488.
- [7] J. Li, Z. Luo, Y. Zuo, J. Liu, T. Zhang, P. Tang, J. Arbiol, J. Llorca, A. Cabot, *Appl. Catal., B* **2018**, *234*, 10.
- [8] A. Dubale, Y. Zheng, H. Wang, R. Hübner, Y. Li, J. Yang, J. Zhang, N. Sethi, L. He, Z. Zheng, W. Liu, *Angew. Chem., Int. Ed.* **2020**, *132*, 13995.
- [9] B. Zhao, J. Liu, C. Xu, R. Feng, P. Sui, L. Wang, J. Zhang, J. L. Luo, X. Z. Fu, *Adv. Funct. Mater.* **2020**, *31*, 2008812.
- [10] C. Liu, W. Zhou, J. Zhang, Z. Chen, S. Liu, Y. Zhang, J. Yang, L. Xu, W. Hu, Y. Chen, Y. Deng, *Adv. Energy Mater.* **2020**, *10*, 2001397.
- [11] Y. Wu, J. Tian, S. Liu, B. Li, J. Zhao, L. Ma, D. Li, Y. Lan, X. Bu, *Angew. Chem., Int. Ed.* **2019**, *58*, 12185.
- [12] H. Huang, C. Yu, X. Han, H. Huang, Q. Wei, W. Guo, Z. Wang, J. Qiu, *Energy Environ. Sci.* **2020**, *13*, 4990.
- [13] X. Zhao, P. Pachfule, A. Thomas, *Chem. Soc. Rev.* **2021**, *50*, 6871.
- [14] J. Du, F. Li, L.-C. Sun, *Chem. Soc. Rev.* **2021**, *50*, 2663.
- [15] S. Royuela, E. Martinez-Perinan, M. P. Arrieta, J. I. Martinez, M. M. Ramos, F. Zamora, E. Lorenzo, J. L. Segura, *Chem. Commun.* **2020**, *56*, 1267.
- [16] H. Zhu, M. Lu, Y. Wang, S. Yao, M. Zhang, Y. Kan, J. Liu, Y. Chen, S. Liu, Y. Lan, *Nat. Commun.* **2020**, *11*, 497.
- [17] J. Liu, D. Yang, Y. Zhou, G. Zhang, G. Xing, Y. Liu, Y. Ma, O. Terasaki, S. Yang, L. Chen, *Angew. Chem., Int. Ed.* **2021**, *60*, 14473.
- [18] T. Li, W.-D. Zhang, Y. Liu, Y. Li, C. Cheng, H. Zhu, X. Yan, Z. Li, Z.-G. Gu, *J. Mater. Chem. A* **2019**, *7*, 19676.
- [19] B. P. Biswal, S. Chandra, S. Kandambeth, B. Lukose, T. Heine, R. Banerjee, *J. Am. Chem. Soc.* **2013**, *135*, 5328.
- [20] L. Wang, Y. Ni, X. Hou, L. Chen, F. Li, J. Chen, *Angew. Chem., Int. Ed.* **2020**, *59*, 22126.

- [21] Y. Liao, H. Wang, M. Zhu, A. Thomas, *Adv. Mater.* **2018**, *30*, 1705710.
- [22] F. L. Li, Q. Shao, X. Huang, J. P. Lang, *Angew. Chem., Int. Ed.* **2018**, *57*, 1888.
- [23] J. Jin, Y. Zheng, S.-z. Huang, P.-p. Sun, N. Srikanth, L. B. Kong, Q. Yan, K. Zhou, *J. Mater. Chem. A* **2019**, *7*, 783.
- [24] H. Yang, Q. Lin, C. Zhang, X. Yu, Z. Cheng, G. Li, Q. Hu, X. Ren, Q. Zhang, J. Liu, C. He, *Nat. Commun.* **2020**, *11*, 593.
- [25] A. J. Bard, L. R. Faulkner, *Electrochemical Methods Fundamentals and Applications*, John Wiley & Sons, Inc., New York **2001**, pp. 1–850.
- [26] D. Chen, S. D. Minteer, *J. Power Sources* **2015**, *284*, 27.
- [27] G. Chen, Y. Luo, L. Ding, H. Wang, *ACS Catal.* **2018**, *8*, 526.