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Darlene E. Allred

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To the Graduate Council:

I am submitting herewith a thesis written by Darlene E. Allred entitled "The potential long-term influence of fly ash and lime-stabilized sewage sludge on mine spoil reclamation : simulated weathering." I have examined the final electronic copy of this thesis for form and content and recommend that it be accepted in partial fulfillment of the requirements for the degree of Master of Science, with a major in Plant, Soil and Environmental Sciences.

M.E. Essington, Major Professor

We have read this thesis and recommend its acceptance:

J.T. Ammons, J.E. Foss, M. Mullen

Accepted for the Council:

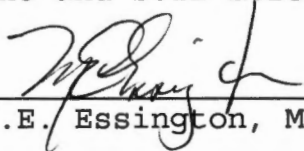
Carolyn R. Hodges

Vice Provost and Dean of the Graduate School

(Original signatures are on file with official student records.)

To the Graduate Council:

I am submitting herewith a thesis written by Darlene E. Allred entitled "The Potential Long-Term Influence of Fly Ash and Lime-Stabilized Sewage Sludge on Mine Spoil Reclamation." I have examined the final copy of this thesis for form and content and recommend that it be accepted in partial fulfillment of the requirements for the Master of Science, with a major in Plant and Soil Science.



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J. T. Ammons



J. E. Foss



M. Mullen

Accepted for the council:



Associate Vice Chancellor and
Dean of the Graduate School

**The Potential Long-Term Influence of Fly Ash and Lime-
Stabilized Sewage Sludge on Mine Spoil Reclamation:
Simulated Weathering**

A Thesis

Presented for

The Master of Science

Degree

The University of Tennessee, Knoxville

Darlene E. Allred

May 1998

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Dedication

This thesis is lovingly dedicated to my grandparents, Mr. and Mrs. Van and Ruby Garvin, Mrs. Lillian Allred, and the late Mr. Thomas Leon Allred, whose love and understanding will always be cherished.

Abstract

Surface mining allows pyrite to be brought to the surface, where it is readily oxidized, producing acid mine drainage. A feasible alternative for the reclamation of mine spoil is the utilization of lime-stabilized sewage sludge (LSS) and fly ash (FA). Lime-stabilized sewage sludge provides essential nutrients to assist in revegetation efforts in an otherwise deficient environment, provides organic matter to improve spoil physical characteristics, and provides a mechanism for disposal. Using a neutral FA in co-application with LSS can enhance mine spoil nutrient levels and physical characteristics. The University of Tennessee, Tennessee Valley Authority, and Electric Power Research Institute initiated a field study in 1994 to address the feasibility of LSS and FA co-utilization for mine spoil reclamation. To complement the field study, simulated laboratory weathering of unamended and amended mine spoil was initiated to assess the potential long-term influence of LSS and FA on mine spoil systems.

The total elemental content of the FA, LSS, and mine spoil was determined by employing both HNO_3 extraction and total digestion (aqua regia/HF). The percent recovery of elements by HNO_3 digestion was determined by dividing the HNO_3 -extractable concentrations by the total elemental content determined using the aqua regia/HF method.

Variations in HNO_3 -extractability of metals in the mine spoil, LSS, and FA was a function of element speciation in the mineral phase. In general, HNO_3 digestion was more efficient at extracting Al, Ca, Mg, Cr, Cu, Pb, and Zn from LSS; K and Na from the FA; and Fe, Mg, P, Co, and Mn from mine spoil. Further, HNO_3 -extractability was not a function of total concentrations and the HNO_3 digestion method did not yield a constant percent extractability that could be uniformly applied, irrespective of material.

The chemical analyses of amended mine spoil leachates showed a neutralization of acidity generated by pyrite oxidation. Slight pH and nitrate fluctuations were observed during the later weathering cycles, which may have been due to increased microbial activity. Also, the initially high electrical conductivity (EC) of the mine spoil leachates was increased further upon LSS application and did not vary with FA rate. The EC of all mine spoil leachates decreased with weathering and mirrored the behavior of Ca and SO_4 . The dissolution and mobilization of Al, Fe, K, Cu, Mn, Ni, and Zn decreased upon LSS application, irrespective of FA rate.

A sequential-selective dissolution (SSD) procedure partitioned elements (Ba, Co, Cr, Cu, Mn, Ni, Pb, Sr, and Zn) into the following operationally-defined pools: (1) soluble-exchangeable, (2) adsorbed, (3) organic, (4) carbonate, (5) sulfide, and (6) residual. The speciation of Ba increased in the soluble-exchangeable and carbonate

fractions with increasing FA rate. Strontium increased in the soluble-exchangeable and sulfide fractions with increasing FA rate. Weathering tended to shift Ba to the carbonate fraction and Sr showed no significant shifts to any other phase. Chromium found in the residual fraction significantly increased upon LSS application. Weathering of the LSS-amended mine spoil increased Cr found in the adsorbed fraction, yet maintained predominance in the residual fraction. Chromium solid-phase speciation was not impacted by FA application or weathering. Although primarily found in the residual fraction, Co showed a significant decrease in the soluble-exchangeable form upon LSS and FA application in the weathered material. In general, FA did not impact Co solid-phase speciation. However, LSS application decreased Co found in the soluble-exchangeable fraction. Copper speciation decreased in the soluble-exchangeable and residual fractions upon LSS application with an increase in the carbonate fraction. Primarily, weathering shifted Cu into the residual forms. Fly ash rate had no influence on Cu speciation for the weathered or unweathered mine spoil. In general, increasing FA rate increased Ni found in the residual fraction. Weathering shifted Ni into the carbonate and sulfide forms which decreased with FA application rate. However, Mn was primarily found in the soluble-exchangeable and residual fractions, irrespective of amendment. Weathering shifted Mn

into the carbonate form in the amended mine spoil upon LSS application and into the residual form for the unamended mine spoil. Upon LSS application, Pb decreased in the soluble-exchangeable fraction. However, when weathered, amendment did not influence Pb solid-phase speciation. In general, FA rate increased the carbonate form of Zn in the unweathered material. However, weathering shifted Zn from the soluble-exchangeable and sulfide forms into the carbonate fraction, which was influenced by the LSS. In general, the impact of FA co-application with LSS was element specific with respect to the solid-phase speciation of elements.

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Chapter I

Introduction

Surface mining for Eastern U.S. coal allows pyrite to be brought to the surface, where it is readily oxidized, producing acidic conditions. The reduced pH results in the release of aluminum and iron which hydrolyze, further reducing solution pH and resulting in the solubilization and mobilization of metal cations. High acidity also inhibits the establishment of vegetation, thereby increasing runoff and erosion. Reclamation of these mining areas is one method of controlling the production of acidity through pyrite oxidation.

Many studies have shown that the application of lime-stabilized sewage sludge (LSS) increases the pH of spoil material, favoring revegetation and drastically reducing erosion. Lime stabilized sewage sludge is a reservoir of nutrients, organic matter, and beneficial microorganisms which are typically needed in mine spoil systems. Further, potential acidity of the mine spoil is neutralized by the alkalinity of the LSS. Increasing the pH also minimizes the activity of *Thiobacillus* (sp.) bacteria, resulting in reduced pyrite oxidation and acid production. Moreover, increasing the pH of the mine spoil material reduces the solubilization and subsequent hydrolysis of aluminum and other metals. For these reasons, LSS is an

effective and viable option for reclaiming surface mining areas.

Many coal-fired power generation plants use coal obtained from surface mining operations. Fly ash (FA) is a by-product of coal combustion. There are approximately 1.8 million tons of FA available in Tennessee. Fly ash is mainly disposed of in slurry ponds or in specified landfills. Fly ash is also utilized for construction purposes (e.g., filler in concrete). The utilization of FA in mine spoil reclamation has been shown to enhance plant micronutrient availability and to enhance the physical characteristics of the spoil. While many studies have addressed the chemical effects and attributes of utilizing alkaline FA, very few have identified the impact of neutral and acidic FA. Further, studies that examine the co-utilization of neutral FA with LSS are basically nonexistent.

The application of large quantities of FA during the reclamation of mine spoil material may be a mechanism for FA disposal. The University of Tennessee, the Tennessee Valley Authority (TVA), and the Electric Power Research Institute initiated a field study in Caryville, TN to evaluate co-utilization of FA and LSS for mine spoil reclamation. This field study, initiated in 1994, will take many years to evaluate the potential long-term chemical impact of FA on the mine spoil environment. Yet, the potential long-term

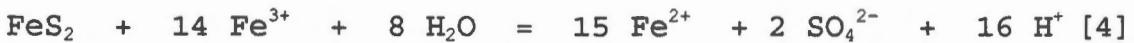
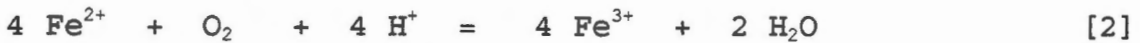
impact of co-utilization of FA and LSS can be evaluated by accelerating the weathering process through the use of simulated laboratory weathering techniques. Alternating wetting-drying cycles in the laboratory environment hastens weathering, allowing for an evaluation of potential long-term impact of LSS and FA applications to mine spoil. The objectives of this study are to (1) to examine the HNO₃-extractable elemental content of LSS, FA, and mine spoil for comparison to total elemental content, (2) to evaluate the leachate chemistry of LSS and FA amended mine spoil during simulated laboratory weathering, and (3) to examine the solid-phase speciation of elements in LSS and FA amended, weathered and unweathered mine spoil.

Chapter II

Literature Review

The neutralization of acidity produced by the oxidation of sulfides is key in reclaiming mine spoil areas.

Typically, acid mine spoil leachates result from the oxidation of pyrite, which may be characterized by the following reactions (Stumm and Morgan, 1981):



The microbially mediated (*T. thiooxidans*) oxidation of sulfide in pyrite by dissolved oxygen to form sulfate [Reaction 1] releases ferrous iron and acidity. The ferrous iron can then undergo microbially mediated (*T. ferrooxidans*) oxidation by consuming dissolved oxygen and protons to produce ferric iron [Reaction 2], which may hydrolyze to form ferric hydroxide and acidity [Reaction 3]. Overall, the oxidation of pyrite to form ferric hydroxide releases four moles of proton for every one mole of pyrite oxidized. Ferric iron may also oxidize the sulfide in pyrite [Reaction 4] releasing ferrous iron which may reenter the cycle through reaction 2.

Sullivan et al. (1988) suggested that secondary mineral formation may also control acid generation and iron/aluminum

solubility. Sullivan et al. (1988) determined that as leaching time increased, pH decreased, and metal solubility increased. They proposed that if secondary minerals control aluminum and iron activity, thermodynamic constants can be used to predict long-term water quality. If the oxidation of FeS_2 occurs and secondary minerals form, iron and aluminum activity can be predicted by the following:

1. If $\text{pH} > 6.0$ and $\text{pe}+\text{pH} < 11.0$ (moderately reducing to reducing environment), sulfate activity is not predicted to influence aluminum and iron activity. Aluminum is controlled by $\text{Al}(\text{OH})_3$ (s) and iron is controlled by $\text{Fe}(\text{OH})_3$.
2. If $\text{pH} < 6.0$ but $\text{pH} > 3.0$ and $\text{pe}+\text{pH} < 11.0$, then aluminum and iron activities are controlled by sulfate secondary mineral formation (e.g., AlOHSO_4 , FeOHSO_4).
3. If $\text{pH} < 3.0$ and $\text{pe}+\text{pH} > 11.0$ (oxidizing environment), then iron will be controlled by ferrous sulfate mineral formation.

For all three conditions, equilibrium pH is determined by iron and aluminum hydrolysis and secondary mineral formation. These results reflect a reaction time of one week by the humidity cell technique before leachate analysis (Sullivan et al., 1988).

Overburden in mining areas of the U.S. generally consists of shales or sandstones which have low base content

and are acidic upon weathering. There are two distinct groups, according to geographical location: (1) western U.S. coal fields, and (2) eastern U.S. coal fields. Western U.S. coal fields generally generate low sulfur coal. Eastern U.S. coal fields produce high sulfur coal containing extensive pyritic deposits (Evangelou, 1995). During the regrading process of mining, these overburden and coal refuse materials are typically deposited at the surface. When this material is weathered and the pyritic materials are oxidized, acidity can be generated before vegetation is restored (Swaine and Goodarzi, 1995).

The acidity generated through pyrite oxidation allows for the mobilization and solubilization of metals such as aluminum, iron, and manganese. Solubilization of these metals contributes to active acidity via hydrolysis reactions and can generate phytotoxic concentrations that minimize revegetation efforts (Swaine and Goodarzi, 1995). However, overburden that contains pyrite has a greater residual or potential acidity relative to active acidity (Doolittle et al., 1993). Pyrite oxidation may continue to produce sulfuric acid for many years after mining operations have ceased (Dent, 1992). This potential acidity, the acidity that is continually generated, must be neutralized for successful reclamation (Evangelou, 1995; McIntosh and Kriesel, 1992).

In most instances, acidity and infertility result in

poor revegetation (Olyer, 1988). The pH of pyritic spoil material normally ranges between 2.5 and 4.0. However, acidity alone does not inhibit plant growth (Sopper and Kerr, 1979; Dent, 1992). Phytotoxic concentrations of soluble aluminum, iron, and manganese in spoil materials may also contribute to poor revegetation (Sopper and Kerr 1979; Olyer, 1988; Dent, 1992). Also, phosphorus and nitrogen are nutrients that are typically limiting in spoil materials further restricting revegetation efforts (Walker, et al., 1996; Evangelou, 1995; Smith, 1996).

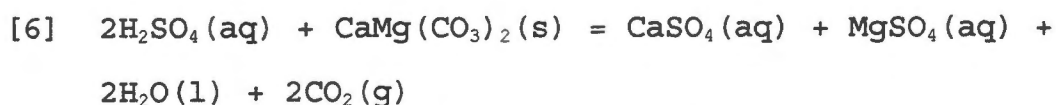
Reclamation Alternatives

Several mechanisms for the reclamation of mine spoils have been proposed and utilized. These include the use of liming agents (such as alkaline earth carbonates) to neutralize potential acidity, bactericides to minimize microbially mediated iron and sulfide oxidation, wetlands to treat acid mine drainage, and organic waste to provide neutralizing capacity, plant nutrients, and physical conditioning (Evangelou, 1995).

The three forms of acidity present in pyritic systems are pore solution acidity, surface acidity, and potential acidity. Pore solution acidity refers to the acidity present in percolating water. Surface acidity refers to the exchangeable acidity on mineral surfaces. Combined, pore solution acidity and surface acidity are defined as active

acidity. Potential acidity is the acidity that can be generated by the spoil material through the oxidation of pyritic materials (Evangelou, 1995).

Neutralization of active and potential acidity can be achieved by using carbonate sources (e.g., calcite, dolomite). Acid-producing potential is defined as the difference between the potential acidity and the neutralizing potential of a material. If the acid-producing potential is negative, acidity will be generated. Conversely, if the acid-producing potential is positive, acidity will not be generated. In mine spoils, potential acidity can be extremely high; thus, neutralization potential of the liming source must equal the acid-producing potential to attain neutralization. Acid neutralization is achieved according to the following reactions:



Calcite and dolomite also have the ability to neutralize the acidity generated by iron and aluminum hydrolysis, resulting in the precipitation of hydroxides (e.g., gibbsite, goethite, ferrihydrate) and basic sulfates (e.g., jurbanite, jarosite). However, the neutralization capacity of the liming materials is short-lived due to the continual amount and rate of acid production from the potential acidity fraction and the relatively high mobility of the alkalinity

(Evangelou, 1995).

Rock phosphate or refined apatite rock $[\text{Ca}_5(\text{PO}_4)_3(\text{F}, \text{Cl}, \text{OH})]$ has also been used to neutralize the acidity generated by pyrite oxidation (Stiller et al., 1985). Rock phosphate may control pyrite oxidation by acting as a weak base in solution and enhancing precipitation of iron phosphate minerals (e.g., strengite, vivanite) (Huang and Evangelou, 1992 and 1994). Yet, the beneficial aspects of applying rock phosphate are short-lived due to the accumulation of iron oxide coatings on the rock phosphate, preventing the further release of phosphate (Stiller et al. 1986; Loomis and Hood, 1984; Evangelou et al., 1992).

Bactericides have been used in recent years to mitigate acid mine drainage. Detergents, organic acids, and food preservatives have been shown to impeded bacterial growth. (Erickson and Ladwig, 1985; Dugan, 1987). These surfactants have been shown to reduce acid mine drainage by 60 to 90 percent (Evangelou, 1995). The two main limitations with using these amendments are (1) the surfactants must be applied repeatedly to prevent microbial repopulation due to high mobility, and (2) adsorption of the surfactant to mineral phases present might limit surfactant availability to the bacteria (Erickson and Ladwig, 1985; Evangelou, 1995).

The creation of wetlands to control acid mine drainage

is also an alternative for acid mine reclamation. The potential removal of metals from the aqueous phase in wetland systems is accomplished by the formation of sparingly-soluble metal sulfides, iron oxyhydroxides, plant uptake, and organic complexation. Because plant uptake and organic complexation is relatively finite, the precipitation of metals as oxides or sulfides is the most effective means of long-term metal removal. Staub and Cohen (1992) and Hammack and Edenborn (1991) studied the removal of Co, Cu, Cd, Ni, Pb, and Zn by sulfide precipitation mediated by sulfate-reducing bacteria. Because this process requires an anaerobic environment, this alternative has been scrutinized due to the uncertainty in maintaining saturation thus limiting the process (Wildeman, 1991).

The application of organic wastes is a beneficial treatment alternative for the reclamation of acidic mine spoils. The growth of beneficial microorganisms may limit *Thiobacillus* (sp.) activity by competing for oxygen. Complexation of Fe(III) and pyritic iron(II) by humates may also limit oxidation of pyritic material by limiting *Thiobacillus* (sp.) catalyzation. Luther et al. (1992) identified that organic wastes might actually contribute to pyritic oxidation by Fe(II) complexation by oxygen containing ligands, thereby acting as the ligating atom. Solubilization of Fe(OH)₃ may promote pyrite oxidation by forming Fe(III)-carboxylate complexes (e.g., Fe(III)-

catechol and $\text{Fe(III)[C}_6\text{H}_4\text{O}_2]^+$). These complexes may adsorb onto pyritic surfaces acting as electron acceptors (Luther et al., 1992). This can be remediated by utilizing well-decomposed organic wastes, composted or treated wastes (Loomis and Hood, 1984). However, a limitation of use is that Fe(III)-organic complexes tend to decompose over time by electron transfer from the organic fraction yielding soluble Fe(II) and CO_2 . Thus, soluble Fe(II) is able to reenter the oxidation process resulting in acid production. To remediate this problem, additional application of waste material might be needed (Evangelou, 1995).

Lime-Stabilized Sewage Sludge

Amendments have been used to neutralize acidity generated through pyrite oxidation and to impart fertility for successful revegetation. Liming agents, such as dolomite and calcite alone, lack organic matter and other nutrients that promote revegetation of mine spoils. Lime-stabilized sewage sludge (LSS) is a source of organic matter, essential nutrients, microorganisms, and alkalinity (Sloan and Basta, 1995). For these reasons, LSS has been used as an aid in successful revegetation efforts in mine spoil reclamation (Walker, 1993; Mathias et al. 1979, Griebel, et al., 1979).

Lime-stabilized sewage sludge can contain up to 25% organic carbon to aid in increased infiltration, increased water-holding capacity, and complexation of toxic metals

(Sloan and Basta, 1995). Organic matter in mine spoils is necessary to establish vegetation because most mining areas lack soil properties necessary to maintain vegetation (Dent, 1992). Also, organic matter enhances soil physical structure by enhancing aggregation of fragmented conglomerants by glues produced by microbial activity (Metting, 1993; McIntosh and Kriesel, 1992; Smith, 1991).

Neutralization of the potential acidity in spoil material by the application of alkaline amendments is commonly employed in mine spoil revegetation efforts (Sullivan et al., 1986; Mays and Bengtson, 1978). The high CaCO_3 equivalent of LSS can neutralize acidity in mine spoil solutions generated by the oxidation of pyrite. Alkaline sewage sludge may also increase solution pH which, in turn, can facilitate the precipitation of iron and aluminum (Sloan and Basta, 1995; Cavallaro et al., 1993; Kardos et al., 1979). Further, aluminum toxicity can be minimized by application of LSS by means of precipitation as well as organic complexation in mine spoil systems (Cavallaro et al., 1993). Regulations existing before 1993 prevented direct land application on acid soils because of pH requirements prior to application. New EPA regulations are not contingent on pH maintenance and allow for the possibility of LSS to be a potential liming material (Sloan and Basta, 1995).

By increasing the pH of the mine spoil system using a

liming agent, microbial catalyzation of pyrite oxidation can be dramatically reduced, subsequently decreasing the amount of acidity generated (Doolittle et al., 1993). Reduction in microbial catalyzation of pyritic oxidation significantly reduces acidity produced by maintaining a pH above 4.0 such that *Thiobacillus* (sp.) activity diminishes (Burt and Caruccio, 1986).

Sopper and Seaker (1983) noted that improvements in water-holding capacity and percent water-stable aggregates were attributed to increased loading rates of LSS amendments on the mine spoil studied. An increased availability of N and P as well as decreased acidity in acid-producing mine spoil materials with the utilization of LSS has also been observed (Walker, 1993).

Studies have also shown many beneficial effects of LSS application for reclamation including enhancement of beneficial microorganisms (Fresquez and Linderman, 1982) and increased vegetation and nutrient supply (Mathias et al., 1979; Berg, 1983).

Fly Ash

Fly ash has been studied as a possible amendment to aid in the reclamation of mine spoils (Stehouwer et al., 1995; Haering and Daniels, 1991; Singh et al., 1982; Capp, 1978). Fly ash production in Tennessee was 2.3 million tons in 1991, with 1.8 million tons available for disposal or

utilization (McIntosh and Kriesel, 1992). Disposal alternatives for FA include ponding, placement in landfills, utilization in construction, and possibly utilization in acid mine spoil reclamation. Fly ash typically consists of small glass-like spheres ranging from 0.01 μm to 100 μm in diameter. Because FA is composed of small particles, it generally has a large surface area to mass ratio which can enhance infiltration, nutrient retention, and water holding capacity of the mine spoil (McIntosh and Kriesel, 1992).

Fly ash particles may also be coated with metal oxides from the combustion process (Mattigod et al., 1990). These coatings could alter aquatic environments by releasing metals in low pH systems (Theis and Wirth, 1977). However, when FA is used in co-application with LSS, mobilization of metals may be reduced due to the neutral to alkaline pH and high sorption capacity of organic matter present in the LSS (McIntosh and Kriesel, 1992). Keefer et al. (1979) studied the effects of applied organics and fly ash to mine spoil. Legume growth increased on the co-amended mine spoil relative to either material alone.

Leaching of arsenic and boron from FA has been of particular concern. Boron can be up to 90% extractable in water alone depending on the FA:water ratio (Pagenkoph and Connolly, 1982). Experiments with acidic FA show that As solubility in water ranges from 2% to 18% of the total As present (Breslin and Duedall, 1983). In addition, As and B

tend to be more soluble in high pH systems, whereas metals such as Pb and Cu are less soluble in high pH systems (McIntosh and Kriesel, 1992). In amended mine spoil systems, the solubilization of both As and B are expected to be less than the water-extractable concentrations due to high affinity for sorption to both organics and minerals present (Phung et al., 1979; McIntosh and Kriesel, 1992).

A bulk of the studies that examine FA utilization focus on alkaline fly ash. However, neutral to acidic fly ash material may also be beneficial in reclamation efforts providing essential macro- and micronutrients. The pH of the spoil system is of importance as stated earlier due to solubilities in low pH systems. When a neutral fly ash is used in coapplication with a liming agent, the pH of the system is favorable for decreased mobility of metals in the system (Capp, 1978).

Coapplication: LSS and FA

Co-mixtures of LSS and FA may enhance mine spoil physical characteristics and nutrient availability and retention. Neutral FA may improve micronutrient availability, promote soil aggregation, and enhance soil physical characteristics (Capp, 1978). For these reasons, amendment co-application can be a beneficial method of FA disposal.

Beaver (1995) examined the addition of coal ash to a

composting mix. In this study, dairy manure/compost was amended with coal ash at rates of 0, 54, 87, 188, and 318 kg MT⁻¹ and the mixtures were monitored for six weeks to determine the impact of coal ash on microbial activity. Microbial activity was not impacted by the coal ash additions. The pH, C:N ratio, and electrical conductivity (EC) of each treatment were determined weekly. The pH and C:N ratios of the plots were relatively constant with values of 8.5 and 27.5, respectively. The EC of the treatments increased with coal ash amendment rate, ranging from 1.25 to 1.95 dS m⁻¹. In growth trials, the biomass of tomato plants grown on dairy manure/coal ash amended soil was greater than soil amended with compost alone. In another growth trial, grain yields increased with increasing coal ash amendments. It was postulated that this yield increase was due to micronutrients (e.g., potassium and magnesium) supplied by the coal ash (Keefer et al., 1979).

Adriano et al. (1982) examined the feasibility of applying a mixture of alkaline fly ash and anaerobically digested sewage sludge in sudan grass production. The rationale for this study was that N and P, which fly ash lacked, was supplied by the sewage sludge fraction of the amendment. The fly ash used was alkaline and predicted to reduce the solubility of metals in the sludge, thereby decreasing metal phytotoxicity. The results of the study show that Cd uptake by sudan grass was reduced with

increasing alkaline fly ash amendments on two California soils with sludge additions containing $30 \text{ mg kg}^{-1} \text{ Cd}$. Crop yields increased with sludge rate but decreased with increasing fly ash rates. This was postulated to be due to the unweathered nature of the fly ash. Unweathered fly ash contributes to toxic metal concentrations due to the ready available metals on the surface of the fly ash particles. Upon weathering of the fly ash, these metals are removed thereby reducing toxic elemental concentrations in leachates. The study was conducted with a weathered fly ash which was found to have no effect on crop yields.

Laboratory Weathering

One method of evaluating the potential long-term effects of amendment applications in a timely manner is by laboratory weathering. The two general types of laboratory weathering are equilibrium and non-equilibrium weathering. Equilibrium weathering is a steady-state evaluation and study of leachate chemistry (Essington, 1991). Non-equilibrium laboratory weathering involves alternating wetting-drying cycles to simulate natural environmental conditions. Two methods that have been employed for non-equilibrium laboratory weathering methods are the soxhlet and the humidity cell techniques (Essington, 1991; Sullivan et al., 1986).

In the soxhlet method, the solids are added to a porous

thimble and placed in the mixing/extraction chamber. Leaching is achieved through the condensation of a heated solvent (H₂O) moving through the sample. In the original soxhlet design, the mixing/extraction chamber was located directly above the boiling solvent, and the temperature of the extracted leachates exceeded normal environmental conditions, limiting microbial activity (Sobek et al., 1982). Since pyritic oxidation is catalyzed by microbial activity, it is important to promote during simulated weathering (Backes et al., 1993).

A modification to the soxhlet design by Sobek et al. (1982) decreased the fluctuation in temperature of the leachate during simulated weathering. This was done by moving the mixing/extraction chamber away from high temperature refluxing solution (Sobek et al., 1982). This modification allows for the introduction and sustained activity of microbes during weathering. However, the rapid leaching environment that the soxhlet provides may not allow for the adequate initiation of microbial activity for pyrite oxidation (Sullivan and Sobek, 1982).

The humidity cell technique involves alternating wetting-drying cycles to simulate natural seasonal fluctuations in soil water content. The humidity cell technique has been used to assess the generation of acidity in both high and low leaching environments (Sullivan et al., 1988). One benefit of using this technique is that

weathering is conducted at ambient temperatures in the laboratory which favors microbial activity (Sullivan and Sobek, 1982). Another benefit of using the humidity cell technique is that the aerobic unsaturated environment allows oxidation to occur in a moist environment for leachate equilibration (Sullivan et al., 1988). One noted pitfall associated with this technique is that it is not as timely as the Soxhlet method, yet it is a relatively reliable method of leachate characterization (Sullivan et al., 1988). In addition, the results from the humidity cell technique need validation with field experiments as well as differences in leaching/equilibration time. This technique provides reproducible results, is relatively short-term and cost-effective, and does not require specialized equipment (Essington, 1991).

Element Solid-Phase Speciation

Due to the concern associated with elevated metal concentrations in sewage sludge amended soils, techniques have been developed to closely monitor the mobility of such metals (Lake et al., 1984). Guidelines that limit land application of sewage sludge were generally developed based on phytotoxic effects founded on plant uptake studies (Lake et al., 1984). Metal distribution in sewage sludge amended soils have been investigated by sequential-selective dissolution (SSD) techniques. Such extraction procedures

can represent only an arbitrary comparison of metal forms and distributions, as the chemical reagents employed lack selectivity (Sterritt and Lester, 1980). The role of reagent use in SSD methods is to discretely remove metals by fractions in order of chemical resistance. Success is not easily achieved because of the lack of validation (Nirel and Morel, 1990; Ma and Uren, 1995). Differences among SSD techniques include reaction time, reagent, reagent sequence, and reagent strength. Although no standard SSD method has been identified, SSD techniques are considered to be more useful than single extractants to determine trace metal distribution in sludge amended soils (Stover et al. 1976).

Sequential-selective dissolution can provide an estimation of the distribution of trace metals in chemical pools (Essington, 1989). The procedure used in this study was developed by Stover et al. (1976) and revised by Lund et al. (1980) for the indirect characterization of trace element solid-phase speciation in sewage sludge and sewage sludge-amended soil. This method partitions trace elements into the following operationally-defined pools: exchangeable, adsorbed, organic, carbonate, and sulfide forms. In addition, trace elements residing in the residual (non-extractable) fraction were computed by the difference between the total metal content and the sum of the extractable metals. The trace element pools in the solid phase were extracted by use of the following reagents: (1)

exchangeable: 0.5 M KNO_3 for 16 hours (2) adsorbed: deionized water for 2 hours repeated 3 times, (3) organic: 0.5 M NaOH for 16 hours, (4) carbonate: 0.05 disodium EDTA for 6 hours, and (5) sulfide: 4 M HNO_3 for 16 hours.

Sequential selective dissolution of trace elements into defined pools can provide information on the distribution of heavy metals in mine spoil systems amended with sewage sludge (Cao et al. 1984). Many trace elements are essential for microbial and plant growth in small amounts, but above optimum concentrations, can become harmful to both plants and animals (Angel and Feagley, 1987). Thus, monitoring trace metals is crucial when introducing potential metal loadings by utilizing amendments. Although many SSD procedures have been scrutinized, SSD is a viable method to monitor the speciation of metals. The majority of heavy metals added by sludge application are acid-extractable (Neuhauser and Hartenstein, 1980) and the results are comparable to stronger acid digestible methods such as nitric acid digestion, perchloric digestion, and hydrofluoric decomposition (Burkitt, et al. 1972; Harrison and Laxen, 1977; Cao et al., 1984). Yet, nitric acid digestion yield results consistently underestimated heavy-metal content in a study conducted by Chang et al. (1983) on sludge-treated soils.

Implemented Field Study

The utilization of sewage sludge has been shown to aid in revegetation efforts for reclamation of mined lands (Sloan and Basta, 1995; Cavallaro et al., 1993; Little et al, 1991). Fly ash has also been studied as a possible amendment for reclamation of mine spoils (Stehouwer et al., 1995; Haering and Daniels, 1991; Singh et al., 1982; Capp, 1978). Keefer et al. (1979) studied the effects of applied organics and fly ash to mine spoil hypothesizing that the materials might complement one another. Co-mixtures of LSS and FA may enhance spoil characteristics and nutrient availability (Taylor and Schuman, 1988). Neutral FA may improve micronutrient availability, hastening of soil aggregation, and enhancement of particle size distribution which may provide a favorable environment for revegetation (Capp, 1978). For these reasons, amendment co-application can be a beneficial method of FA disposal.

To analyze the feasibility of LSS and FA co-application for mine spoil reclamation, The University of Tennessee, the Tennessee Valley Authority (TVA), and the Electric Power Research Institute initiated a field study in Caryville, TN. This field study, initiated in 1994, will take many years to evaluate the long-term chemical impact of FA for reclamation of the mine spoil. Yet, the potential long-term impact of co-utilization of FA and LSS can be evaluated by accelerating the weathering process through the use of

simulated laboratory weathering techniques. Alternating wetting-drying cycles in the laboratory environment hastens weathering, allowing for an evaluation of potential long-term impact of LSS and FA applications to mine spoil.

Objectives

The objectives of this study are as follows:

- (1) to examine the HNO₃-extractable elemental content of LSS, FA, and mine spoil and for comparison to total elemental content.
- (2) to evaluate the leachate chemistry of LSS and FA amended mine spoil during simulated laboratory weathering.
- (3) to examine the solid-phase speciation of trace elements in LSS and FA amended, weathered and unweathered mine spoil.

Chapter III

Materials and Methods

Materials

The mine spoil material was collected from an abandoned coal mine located in the Western Appalachian Plateau/Cumberland Mountains physiological region of East Tennessee (Campbell County). The exact location of the site is $36^{\circ} 17' 51''$ N and $84^{\circ} 15' 48''$ W with an elevation of 930 m. The disturbed site resulted from the surface mining of the Cold Gap Coal Bed of the Cross Mountain Formation in the Middle Pennsylvanian. In 1994, the mine spoil site was subjected to reclamation with LSS and FA amendments as part of a demonstration project for Tennessee Valley Authority. Prior to amendment, surface 30 cm samples of the mine spoil were collected at randomly selected locations at the mine site and composited. The mine spoil samples were transported to Knoxville, TN, allowed to air-dry, thoroughly mixed, and stored in a sealed containers at ambient temperature (22°C to 25°C) until needed. The LSS was obtained from the Knoxville Utility Board's Kuwahee Waste Water Treatment Plant located in Knoxville, TN. This sewage sludge is classified as a Class A, low metal sewage sludge as per the US EPA Part 503 sludge management regulations (Chapter 40 Code of Federal Regulations, Part 503). The LSS was stored in a tightly sealed container at 4°C until needed. The FA

was collected from the Kingston Fossil Plant, operated by TVA and located in Kingston, TN. The FA was dry when delivered and was stored in a sealed container at ambient temperature until needed.

Characterizations

Acid-base accounting yields the net acid or neutralization potential of a material by taking the difference between the neutralization potential and the maximum acid producing potential (Miller and Murray, 1988). Neutralization potential was determined in triplicate by reacting standardized 0.1 M HCl with a known mass of mine spoil, LSS, or FA and boiling the solution for approximately one minute. The mixture was cooled and back titrated with standardized 0.1 M NaOH to determine the amount of HCl consumed. Acid-producing potential was determined by total sulfur analysis with the LECO CR-12 furnace after sulfate removal (Grube et al., 1993; Miller and Murry, 1988). Sulfate removal was accomplished by filtering approximately 50 mL of a 2:3 HCl solution through the material followed by thorough rinsing with distilled, deionized water (US EPA, 1974). The acid-base account is expressed as tons CaCO_3 eq/1000 tons material by using a 31.25 correction factor (assuming that all S is from pyrite or marcasite). Averaged among triplicates, the results of the acid base accounting of the mine spoil and LSS were -12.5 and 151 t

CaCO₃ equivalent per 1000 t dry material, respectively. Thus, the LSS amendment rate, computed on a dry-weight basis, was 88 tons LSS/1000 tons mine spoil (197 Mg ha⁻¹). Also, the FA had an acid-base account of 16.94 tons CaCO₃ eq/1000 tons of material when averaged among triplicates. The moisture content of the LSS was determined to be 12.5% by standard gravimetric evaluation by oven-drying (Gardner, 1986).

The total elemental content of the mine spoil, LSS, FA, and the weathered material was facilitated in triplicate using the microwave-induced fluoroboric acid dissolution technique by Nadkarni (1984) with modification by Ammons et al. (1995). A 200 mg sample (<60 mesh) was pre-treated with 2 mL HF in polyallomer centrifuge tubes, mixed by vortexing, and allowed to react for 16 hours. After pretreatment, 5 mL of *aqua regia* (3:1:1 - HCl:HNO₃:H₂O) was added to each sample and mixed by vortexing. Samples were then heated in a microwave for a 3 minute reaction time at 80% power to hasten the dissolution reaction. After cooling, 1 gram of boric acid was added to each sample to neutralize excess HF present. Samples were mixed by vortexing and heated in the microwave for an additional 10 minutes at 20% power to dissolve the boric acid. While still warm, the samples were mixed by vortexing to facilitate boric acid dissolution. The cooled samples were filtered through Whatman #42 filter paper into 100 mL

volumetrics and brought to volume with deionized water. The chemical analysis was performed by a Thermo Jerrell Ash Model 61 inductively-coupled argon plasma optical emission spectrophotometer (ICP-OES) to determine the concentrations of Al, As, Ba, Ca, Cd, Co, Cr, Cu, Fe, K, Mn, Ni, P, Pb, Se, Sr, and Zn.

The HNO₃-extractable metal content of the mine spoil, LSS, and FA was facilitated by 4 M HNO₃ digestion (Cao et al., 1984). The triplicate digestion tubes contained 3.5 g of sample and 21 mL of 4 M HNO₃ and were placed in a digestion block calibrated to 80°C. The samples were refluxed for 16 hours. After the samples were cooled and brought to 35 mL total volume, the solution was filtered and subjected to ICP-OES analysis to determine Al, As, Ca, Cd, Cr, Cu, Fe, K, P, Pb, Mg, Mn, Ni, Se, and Zn content.

Sequential-selective dissolution analysis was performed on samples from the 2 and 18 week weathering cycles for comparison of metal pool distribution. The procedure used in this study was developed by Stover et al. (1976) and revised by Lund et al. (1980) for the indirect characterization of trace element solid-phase speciation in sewage sludge and sewage sludge-amended soil. This method partitions trace elements into the following operationally-defined chemical pools: exchangeable, adsorbed, organic, carbonate, and sulfide forms. In addition, trace elements residing in the residual (non-extractable) fraction were

computed by the difference between the total metal content and the sum of the extractable metals. Extractions were performed in 50 mL polypropylene centrifuge tubes using triplicate 2 g samples with 25 g the of appropriate reagent. The metal pools of the solid phase were extracted by use of the following reagents: (1) exchangeable: 0.5 M KNO_3 for 16 hours (2) adsorbed: deionized water for 2 hours repeated 3 times, (3) organic: 0.5 M NaOH for 16 hours, (4) carbonate: 0.05 M EDTA for 6 hours, and (5) sulfide: 4 M HNO_3 for 16 hours. The samples were centrifuged between steps for approximately 10 minutes to separate the solid from the extractant. The extracts were analyzed by ICP-OES for As, Ba, Co, Cr, Cu, Mn, Ni, Pb, Se, Sr, and Zn. The mass of a given element extracted by a given reagent was calculated as follows: $\mu\text{g extracted} = (A * 25\text{g}) - (B * C)$ where A is the concentration in $\mu\text{g/g}$ (liquid) extracted, B is the concentration in $\mu\text{g/g}$ (liquid) extracted from the previous step, and C is the mass in grams of extracting solution left over from proceeding extraction step to the current step (Sposito et al., 1982).

The bulk mineralogy of the mine spoil and the FA was determined by powder x-ray diffraction (powder mount) and the clay mineralogy of the mine spoil was determined by x-ray diffraction of Mg-saturated, Mg-glycol, and K-saturated samples. Prior to x-ray diffraction analysis soluble salts and carbonates, Mn oxides and organic matter,

and Fe oxides were removed using standard techniques (Kunze and Dixon, 1986).

A humidity cell technique was used for non-equilibrium laboratory weathering in this study (Essington, 1989). Fifteen 64 X 38 cm polypropylene containers were obtained and weighed. A 500 g sample of mine spoil was placed in each container. The LSS and FA amendment rates, as determined by acid-base accounting were: 0:0, 197:0, 197:197, 197:295, 197:394 Mg ha⁻¹ on a dry-weight basis. Since the objective was to utilize large quantities of FA, FA amendments were 1 X, 1.5 X, and 2 X the calculated amendment rate for the LSS, ignoring the neutralization potential of the FA.

The LSS and FA rates were added to the 500 g spoil sample and thoroughly mixed. A 500 mL volume of deionized water was then added and the amended spoil was allowed to equilibrate for 2 hours. The spoil solution was extracted from the solids by vacuum filtration through Whatman #2 filter paper into a polypropylene Erlenmeyer flask. The leachate was collected in acid-washed polypropylene bottles. The leachates were not allowed to contact glass to maintain the integrity of the leachates for boron analysis. Both the pH and the electrical conductivity (EC) of the leachate were determined immediately after filtration. The remainder of the leachate was capped and stored at 4°C for chemical analysis.

The solid cake collected on the filter paper was returned to the corresponding plastic container and distributed evenly to facilitate drying. The solids were allowed to air-dry for one week, after which time the material was extracted again with a mass of water equal to the mass of the amended spoil material. This weathering technique continued for 18 cycles. Nitrate, sulfate, fluoride, and chloride concentrations in collected leachates were determined by ion chromatography (IC) using Dionex DX-100 ion chromatograph. Leachate Al, B, Ca, Cd, Cu, Fe, K, P, Pb, Mg, Mn, Ni, and Zn concentrations were determined by ICP-OES.

Statistical analyses to compute standard errors the mine spoil, LSS, and FA were performed using MS Excel data analyses VBA tool pack (Microsoft Corporation, 1994). Also, statistical analyses to determine the impact of treatment and weathering cycle on element solid-phase speciation was performed by SAS using the General Linear Model and Student-Newman-Keuls test for determining differences (SAS Institute Inc., 1985).

Chapter IV

Results and Discussion

Clay and Bulk Mineralogy

The clay mineralogy of the mine spoil was determined by x-ray diffraction using standard saturation and treatment methods as detailed in EPA-600/2-78-054 (U.S. EPA, 1978) (Figures 1 and 2). The clay mineralogy of the mine spoil was identified as vermiculite, kaolinite, and mica. The presence of a 1.44 nm peak before and after glycolation indicated the occurrence of vermiculite. Kaolinite was identified from the 0.72 nm and the 0.36 nm peaks that disappeared upon heat treatment at 550 C. Mica was identified by the presence of the 1.0 nm and 0.5 nm peaks in the Mg^{2+} -saturated samples.

The bulk mineralogy of the mine spoil was determined to be primarily composed of quartz. The FA mineralogy was dominated by quartz and mullite ($Al_6Si_2O_{13}$). Mullite was not clearly identified in the LSS/FA-amended mine spoil due to dilution of the FA. The mineralogy of the weathered mine spoil, as determined by x-ray diffraction, was not different from the unweathered mine spoil.

Amendment Characteristics

The total elemental content of the FA, LSS, and mine spoil was determined by employing both HNO_3 extraction and

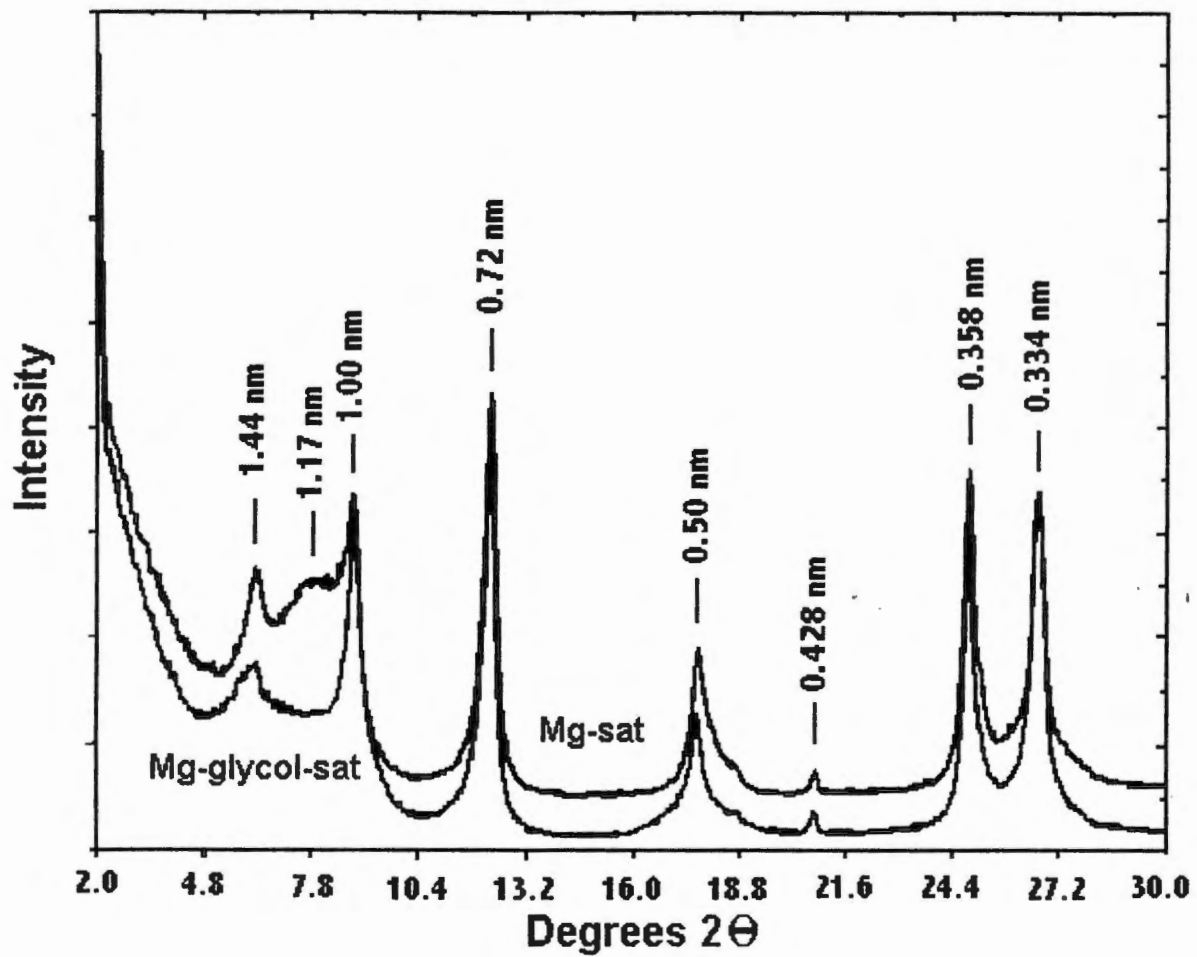


Figure 1. X-ray diffraction patterns of the <2 μm size fraction (Mg-saturated and Mg-glycol saturated) of the Caryville mine spoil.

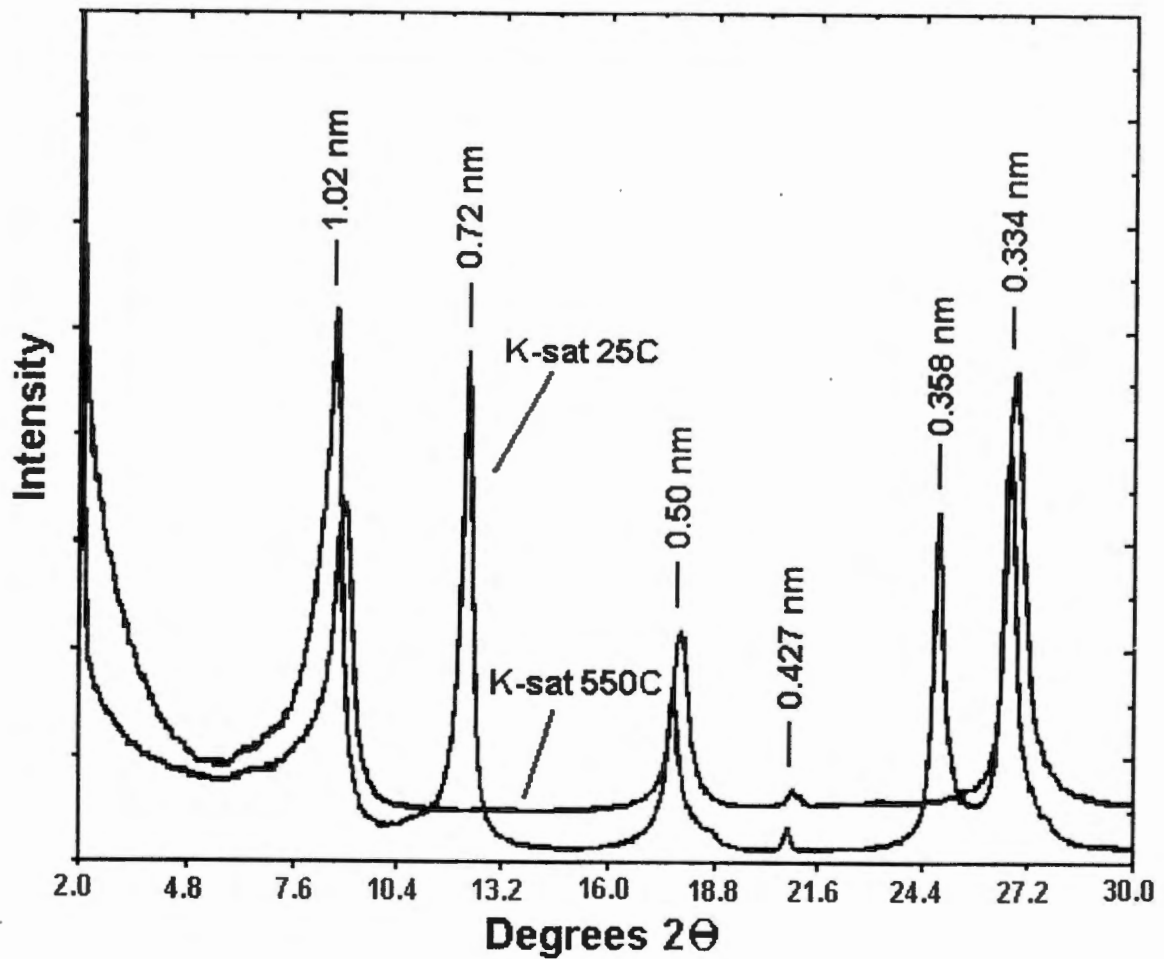


Figure 2. X-ray diffraction patterns of the <2 μm size fraction (K-saturated and K-saturated heat treated at 550°C) of the Caryville mine spoil.

total digestion (aqua regia/HF) (Table 1). The percent recovery of elements by HNO₃ digestion was determined by dividing the HNO₃-extractable concentrations by the total elemental content determined using the aqua regia/HF method.

Although Al content was highest in the FA (106 g kg⁻¹) and lowest in the LSS (13.4 g kg⁻¹), the HNO₃-extractability was greatest for the LSS (37.4%). However, total Fe content and HNO₃-extractability were highest for the mine spoil (49.9 g kg⁻¹ and 100%, respectively) which was expected due to the presence of pyrite. Consequently, Fe content was lowest for the LSS (9.06 g kg⁻¹), yet the HNO₃-extractability was determined to be 82.8%. The HNO₃-extractability of Fe in the FA was determined to be the lowest at 31.8%.

Total Ca content and HNO₃-extractability were greatest for the LSS (58.2 g kg⁻¹ and 91.7%, respectively) and lowest for the mine spoil material (1.86 g kg⁻¹ and 12.7%, respectively). Total K and Mg concentrations were highest in the mine spoil material (9.80 g kg⁻¹ and 11.3 g kg⁻¹, respectively) and lowest for the LSS (3.91 g kg⁻¹ and 7.25 g kg⁻¹, respectively). However, the FA was found to have the greatest K and Mg HNO₃-extractability (73.5% and 21.7%, respectively).

Although total Na content was highest for the mine spoil material (40.3 g kg⁻¹) and lowest for the LSS (27.3 g kg⁻¹), FA had the greatest HNO₃-extractability (2%). Indisputably, Na was in a mineral phase that was resistant

Table 1. Elemental composition of the Kingston fly ash, Knoxville lime-stabilized sewage sludge, and the Caryville mine spoil determined after the HNO₃ extraction and aqua regia/HF dissolution.

Element	Fly ash						Lime-Stabilized Sewage					
	HNO ₃	SE ^{††}	(% T) ^Φ	Total	SE	Range [†]	HNO ₃	SE	(% T)	Total	SE	Range [‡]
g kg ⁻¹												
Al	27.9	2.3	(26.3)	106	1.91	1.0 - 209	5.00	0.07	(37.3)	13.4	1.30	1.0 - 135
Ca	7.40	0.63	(85.0)	8.7	0.10	1.1 - 223	58.2	0.04	(91.7)	63.5	5.45	1.8 - 200
Fe	10.4	1.4	(31.8)	32.7	0.56	10 - 276	9.06	0.04	(83.1)	10.9	0.64	1.0 - 154
K	5.90	0.13	(73.6)	8.02	2.39	1.7 - 67	0.34	0.01	(8.7)	3.91	1.03	0.2 - 39
Mg	2.36	0.11	(21.7)	10.9	5.59	0.4 - 77	1.40	0.03	(19.3)	7.25	2.79	0.3 - 19
Na	0.60	0.03	(6.7)	8.9	3.59	0.1 - 71	0.22	0.002	(0.5)	39.2	1.98	0.1 - 22
P	1.50	0.18	(61.0)	2.46	0.53	----	4.55	0.03	(97.6)	4.66	0.27	5 - 143
mg kg ⁻¹												
Cd	1.36	0.28	(17.7)	7.69	4.29	0.1 - 130	<0.4	NA	(NA)	<0.4	NA	0.6 - 3400
Co	23.8	1.22	(40.7)	58.5	3.91	----	1.61	0.12	(26.7)	6.03	0.90	2 - 2490
Cr	18.2	3.27	(19.6)	92.8	1.17	4 - 900	21.6	0.35	(94.3)	22.9	1.40	tr - 99000
Cu	103	1.08	(53.1)	194	9.32	33 - 2200	63.7	1.4	(84.7)	75.2	7.66	45 - 17000
Mn	82.0	6.06	(35.0)	234	90.9	25 - 3000	247	0.67	(59.5)	415	78.1	32 - 14035
Ni	43.4	6.13	(53.1)	81.8	2.11	2 - 4300	12.5	0.21	(NA)	<10.0	NA	2 - 8000
Pb	60.6	5.32	(26.8)	226	7.00	3 - 2100	19.5	0.62	(35.6)	54.7	13.8	13 - 26000
Zn	106	6.13	(59.9)	177	8.60	14 - 3500	178	0.75	(98.9)	180	7.34	50 - 49000

^{††} Standard error.

^Φ (HNO₃-extractable/Total)* 100

[†] Ranges given for coal fly ash (Mattigod et al., 1990; Eary et al., 1990).

[‡] Ranges given for anaerobically digested sewage sludge (Essington and Mattigod, 1990).

* Mean shale (Bowen, 1979).

Table 1. (continued)

Element	Mine spoil					
	HNO ₃	SE	(% T)	Total	SE	Mean*
----- g kg ⁻¹ -----						
Al	9.00	3.06	(12.4)	72.7	2.68	88
Ca	1.86	0.39	(49.6)	3.75	0.15	16
Fe	49.9	0.93	(100)	44.8	0.77	48
K	1.72	0.81	(17.5)	9.80	1.73	25
Mg	2.22	0.17	(45.5)	4.88	0.10	16
Na	0.08	0.01	(0.4)	17.9	2.13	5.9
P	1.36	0.002	(100)	0.99	0.02	0.7
----- mg kg ⁻¹ -----						
Cd	<0.4	NA	(NA)	<0.4	NA	0.22
Co	11.3	2.69	(65.3)	17.3	2.80	19
Cr	16.4	0.93	(24.9)	65.9	1.19	90
Cu	16.6	1.15	(73.5)	22.6	2.72	39
Mn	278	21.6	(83.7)	332	138.8	850
Ni	26.3	0.30	(NA)	<10.0	NA	68
Pb	27.9	7.57	(20.1)	139	10.43	23
Zn	93.0	9.30	(84.5)	110	27.2	120

†† Standard error.

Φ (HNO₃-extractable/Total)* 100

† Ranges given for coal fly ash (Mattigod et al., 1990; Eary et al., 1990).

‡ Ranges given for anaerobically digested sewage sludge (Essington and Mattigod, 1990).

* Mean shale (Bowen, 1979).

to nitric acid extraction. Conversely, P maintained high HNO_3 -extractability in the mine spoil (100%), LSS (97.6%), and FA (61%).

Although Cd was below the ICP-OES detection limit for both the LSS and the mine spoil, the Cd content in FA was determined to be 7.69 mg kg^{-1} with 17.7% being HNO_3 -extractable. Similarly, Ni concentrations in the LSS and mine spoil were below detectable levels, but was detected to be 81.8 mg kg^{-1} in the FA with an HNO_3 -extractability of 53%.

Copper content was highest in the FA (193.8 mg kg^{-1}) and lowest in the mine spoil material (22.59 mg kg^{-1}). However, the greatest determined HNO_3 -extractability was for the LSS (84.7%) when compared to the mine spoil (73.6%) and FA (53.1%). Although, the Mn content was highest in LSS (415.0 mg kg^{-1}) and lowest in the FA (234.0 mg kg^{-1}) the HNO_3 -extractability was highest for the mine spoil (80%). Further, Zn content was highest in the LSS (180.0 mg kg^{-1}) and lowest in the mine spoil (109.8 mg kg^{-1}) with Zn HNO_3 -extractability being highest for the LSS (98.7%).

The highest Pb concentration was observed in the FA (226.0 mg kg^{-1}) and lowest in the LSS (19.5 mg kg^{-1}). However, the greatest HNO_3 -extractability for Pb was observed for the LSS (35.6%). Also, Co content was highest for the FA (58.5 g kg^{-1}) and lowest for the LSS (6.03 mg kg^{-1}), but the greatest HNO_3 -extractability was observed for the mine spoil (65.3%): Comparatively, Cr content was greatest

for the FA (92.8 mg kg^{-1}), but only 19.6% was HNO_3 -extractable. Chromium concentrations were determined to be the lowest for the LSS (22.9 mg kg^{-1}), yet maintained the highest HNO_3 -extractability (94.3%).

Variations in HNO_3 -extractability of metals in the mine spoil, LSS, and FA was a function of element speciation in the mineral phase. In general, HNO_3 digestion was more efficient at extracting Al, Ca, Mg, Cr, Cu, Pb, and Zn from LSS; K and Na from the FA; and Fe, Mg, P, Co, and Mn from mine spoil. Further, HNO_3 -extractability was not a function of total concentrations and the HNO_3 digestion method did not yield a constant percent extractability that could be uniformly applied, irrespective of material.

Laboratory Weathering

Simulated laboratory weathering was conducted over a period of 18 cycles. A weathering cycle consisted of the following steps: (1) saturation of mine spoil and amended mine spoil with water and equilibration for one hour, (2) separation of leachate from solids by vacuum filtration, (3) analysis of collected leachate, and (4) drying of the solids for one week. Each leaching cycle was designed to simulate water table fluctuation or a rainfall event, to imitate the variation in water saturation, and to accelerate weathering. Leachate pH and EC were determined immediately upon collection. The remaining chemical analyses were completed

within one week of collection by IC and ICP-OES.

The leachate pH increased and maintained pH values near neutral with LSS amendment throughout simulated weathering, except for fluctuations around cycle 15 (Figure 3). Further, leachate pH was not influenced by FA amendment. During cycle 14, leachate pH was lowest for the 0 and 197 Mg ha⁻¹ FA rates. In cycle 15, leachate pH was lowest for the 295 and 394 Mg ha⁻¹ FA rates. The pH of the unamended mine spoil leachate was also slightly depressed during cycles 14 and 15. By cycle 16, the leachate pH of all treatments reverted to pH values observed prior to cycle 14. This fluctuation in leachate pH may have resulted from increased mineralization/nitrification by microorganisms because the leachate pH reduction coincided with increased nitrate levels in the leachates (to be discussed).

The EC in all amended spoil material was higher for cycles 2 through 14 due to the application of the LSS (Figure 4). Fly ash rates did not influence leachate EC. The decrease in leachate EC was due to the removal of soluble salts. The EC during weathering of the unamended mine spoil leachates decreased more rapidly than that of the amended mine spoil material, indicating that the LSS/FA amendments were a source of soluble salts.

Calcium in the amended mine spoil leachates mirrored

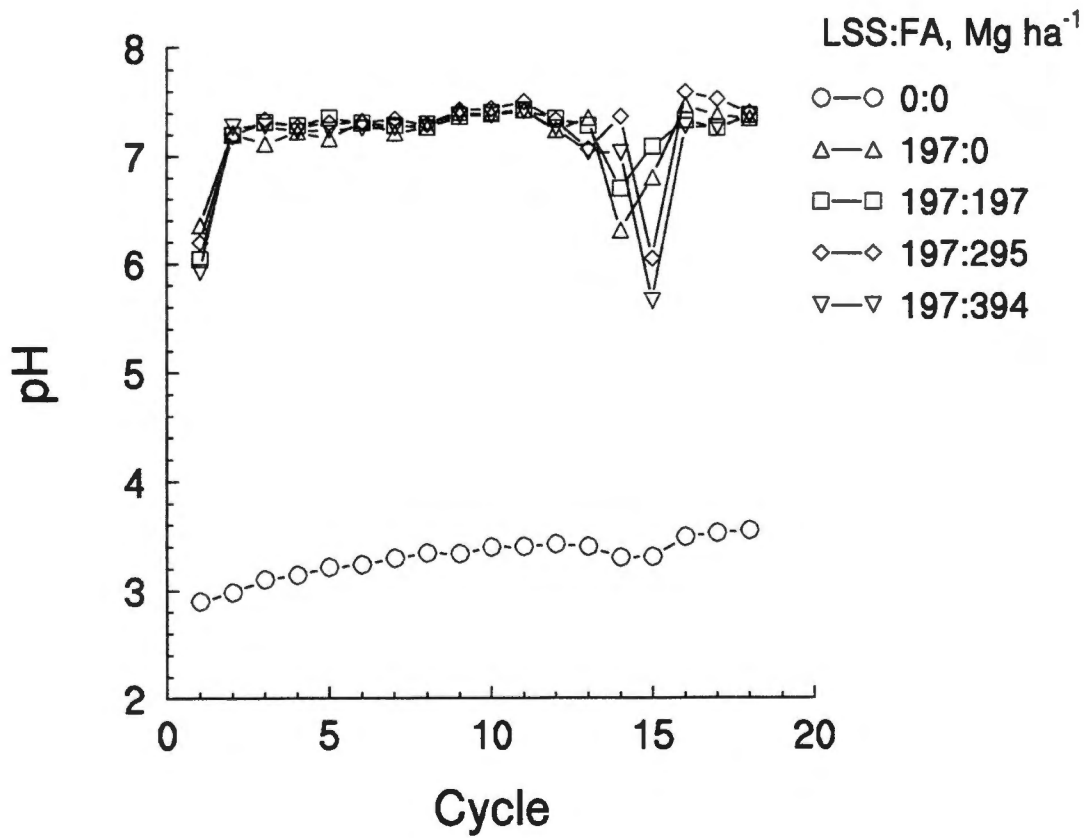


Figure 3. The pH of humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

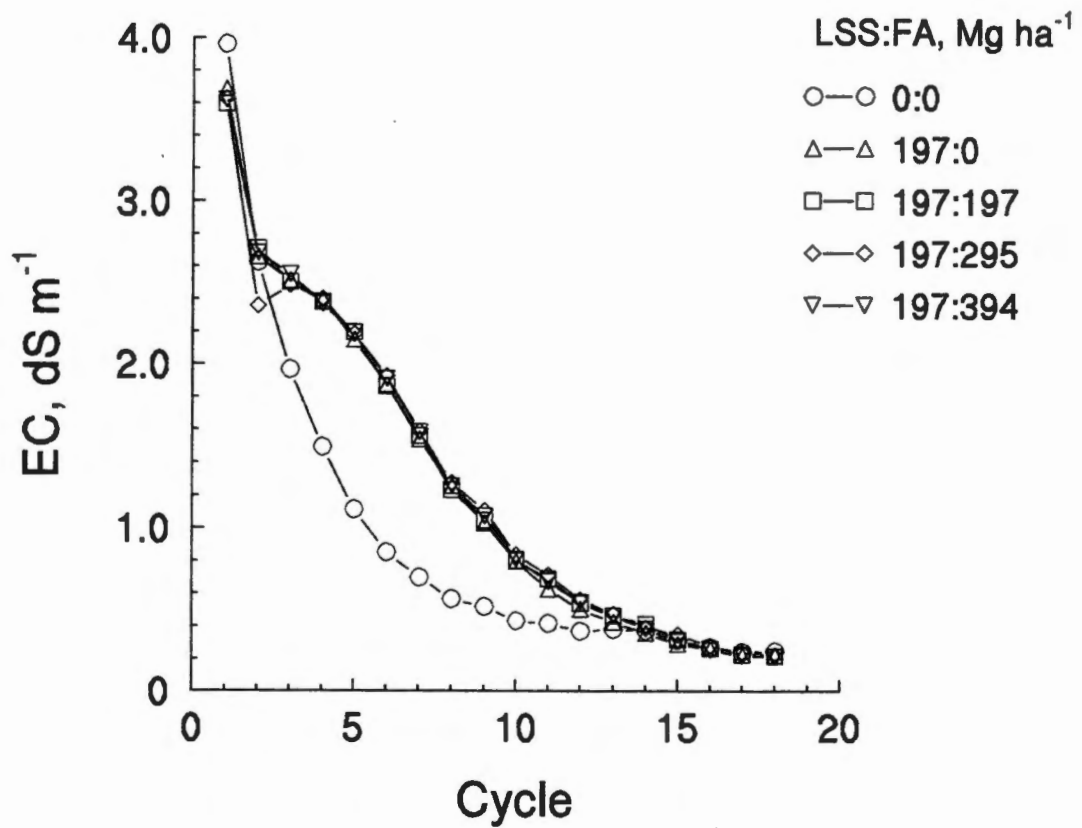


Figure 4. The EC of humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

leachate EC and ranged from a high of 900 mg L⁻¹ during the first weathering cycle to 50 mg L⁻¹ by cycle 18 (Figure 5). Soluble Ca was not influenced by FA amendment rates. The differences in leachate Ca concentration between the unamended mine spoil and amended mine spoil was attributed to LSS application, which supplied the system with additional soluble Ca (Table 1). The Ca concentration in leachate from the unamended mine spoil decreased with weathering. The initial leachate concentration of 450 mg L⁻¹ decreased to 12 mg L⁻¹ by the end of the study. The leaching behavior of Mg was similar to that of Ca shown by the decrease in leachate Mg with weathering (Figure 6). However, while there was a noticeable difference in Mg concentration between the unamended and amended mine spoil leachates, this difference was small in comparison to that observed for Ca.

Leachate sulfate concentrations decreased with weathering for all amended and unamended spoil material (Figure 7). While sulfate concentrations in the unamended and amended mine spoil leachates behaved similarly to that observed for Ca, sulfate concentrations in the unamended spoil leachates were initially greater than in the amended spoil leachates. Further, by the end of the weathering study, all leachates had similar sulfate levels unlike Ca.

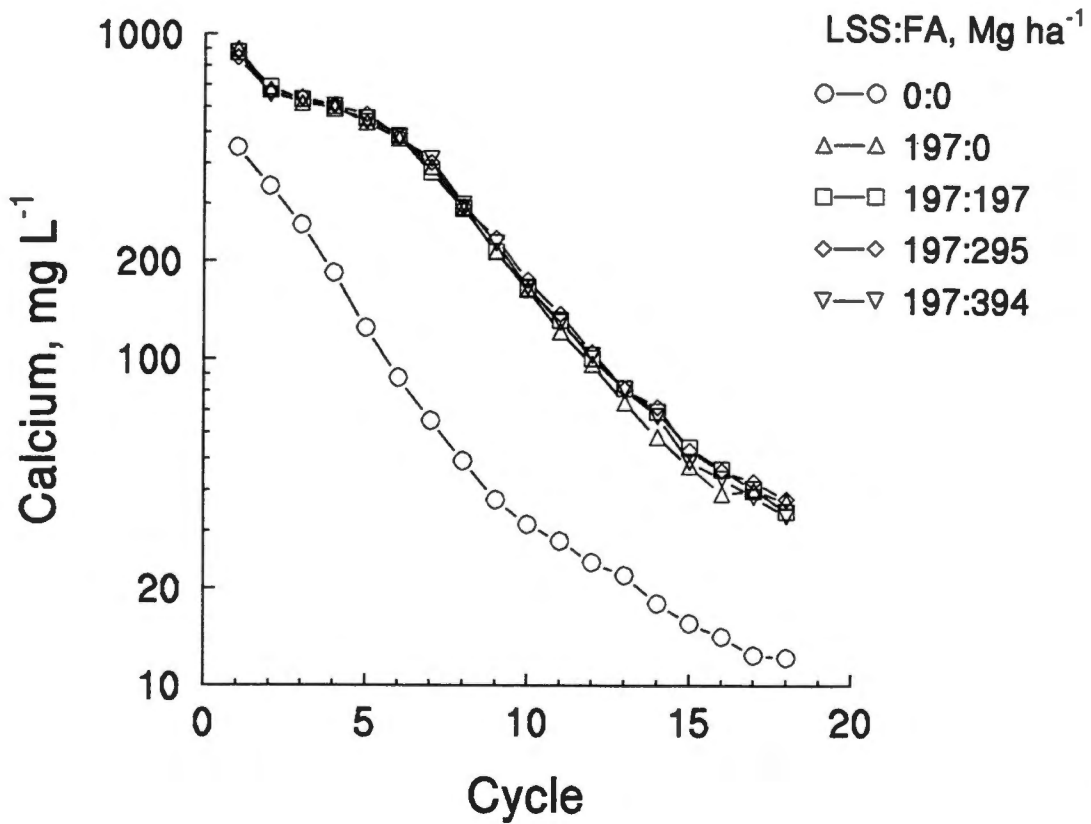


Figure 5. The concentration of calcium in humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

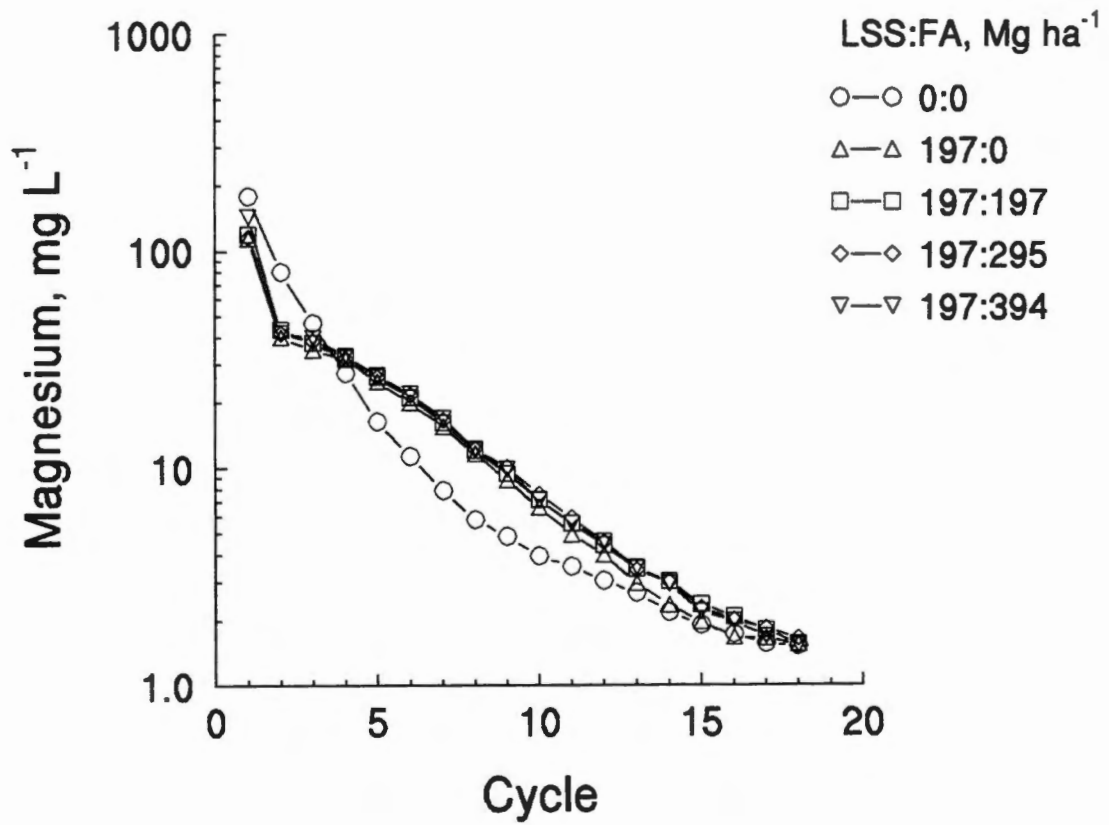


Figure 6. The concentration of magnesium in humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

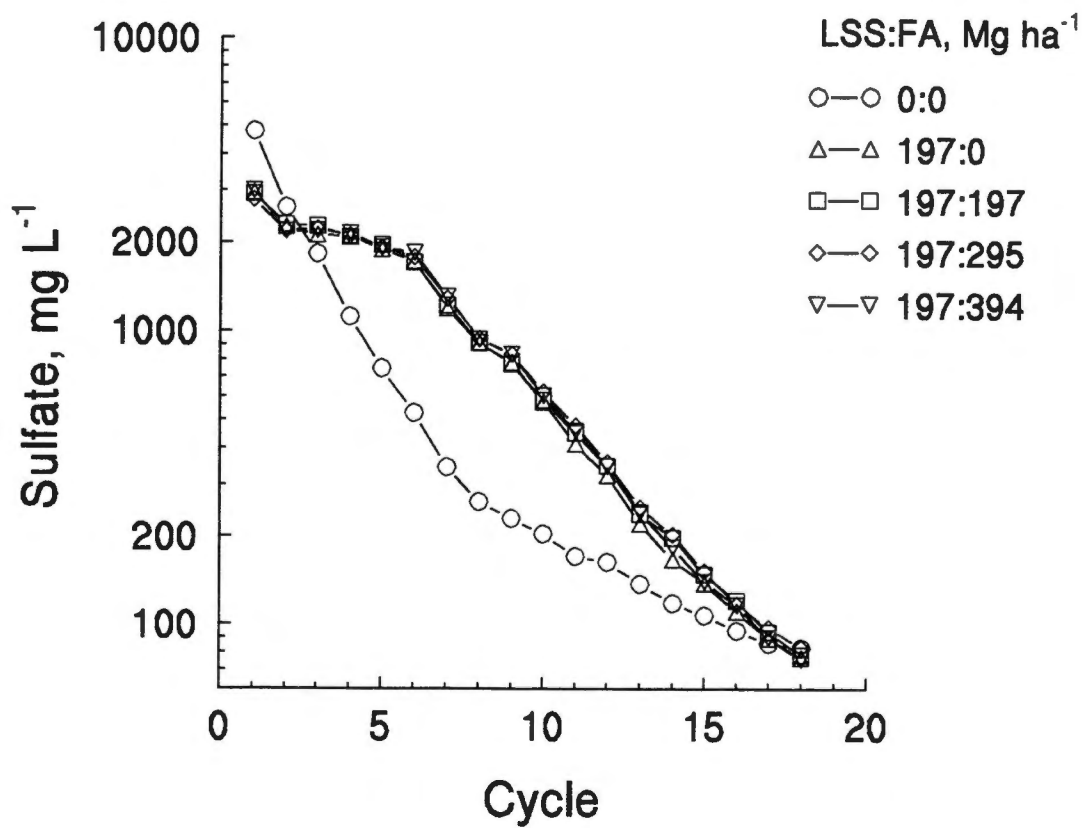


Figure 7. The concentration of sulfate in humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

This was probably due to the microbial catalyzed oxidation of pyrite, which would continue to generate sulfate and acidity. It was evident that pyrite oxidation continued throughout the weathering study as leachate pH remained between 2 and 3 for the unamended mine spoil. The rate of FA application did not influence sulfate concentrations.

Chloride concentrations in amended mine spoil leachate reached 250 mg L^{-1} during the first weathering cycle. Leaching reduced chloride concentrations to approximately 1 mg L^{-1} by the tenth weathering cycle (Figure 8). Clearly, LSS had a pronounced impact on leachate chloride, while FA rates did not influence leachate chloride. Chloride in the unamended mine spoil leachate peaked during the second weathering cycle and was reduced to approximately 0.7 mg L^{-1} by cycle 8.

Fluoride behavior was similar to that of magnesium (Figure 9). Fluoride concentrations peaked at approximately 12 mg kg^{-1} during the first weathering cycle and decreased steadily until cycle 10 when the amended spoil leachate fluoride concentrations were approximately 0.3 mg L^{-1} . The decrease in leachate fluoride concentrations was more rapid in the unamended spoil. The LSS was a source of fluoride and FA rates had no impact on leachate concentrations.

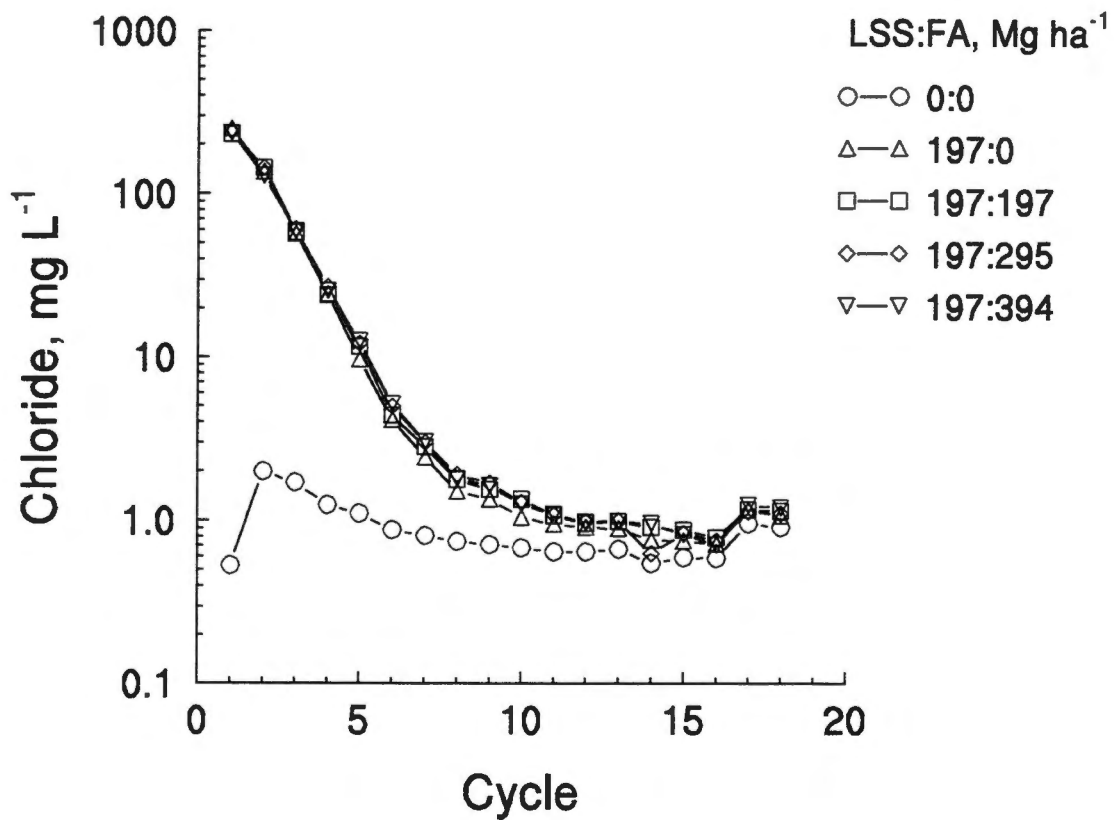


Figure 8. The concentration of chloride in humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

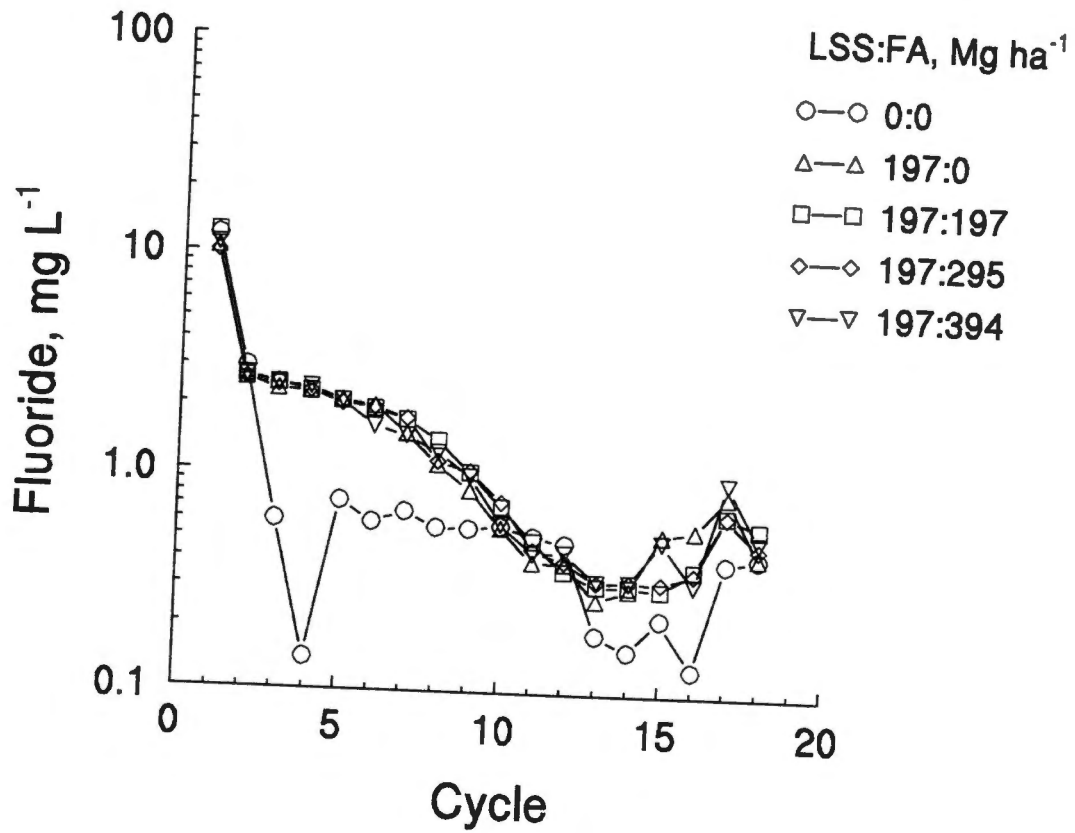


Figure 9. The concentration of fluoride in humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

Phosphorus concentrations in the mine spoil leachate increased with LSS amendment (Figure 10). There also appeared to be a slight trend towards increasing leachate P with FA rate. Concentrations of P in the unamended mine spoil leachates were generally below 0.05 mg L^{-1} throughout the study. However, fluctuations in P correspond to the fluctuations in pH and NO_3 , indicating that P mobility was depressed as pH lowered.

Nitrate leachate concentrations were not influenced by the LSS or FA amendments through the first 12 weathering cycles (Figure 11). Nitrate concentrations smoothly decreased to range between 1 and 2 mg L^{-1} . However, after cycle 11, nitrate concentrations increased to approximately 20 mg L^{-1} . Nitrate leachate concentrations were highest for the 0 and 197 Mg ha^{-1} FA rates (with LSS) during cycle 14. During cycle 15, the unamended mine spoil and the 295 and 394 Mg ha^{-1} FA amended spoil material had the highest leachate concentrations of nitrate. The flush of leachate nitrate that began during cycle 12 coincides with reduced leachate pH levels. Further, the flush of nitrate was not unique to LSS amended mine spoil, as elevated nitrate levels were observed in the unamended mine spoil leachates. Although the reason for this fluctuation was not clear, microbes were likely in sufficient number to oxidize organic-N to nitrate resulting in higher leachate nitrate and acid production.

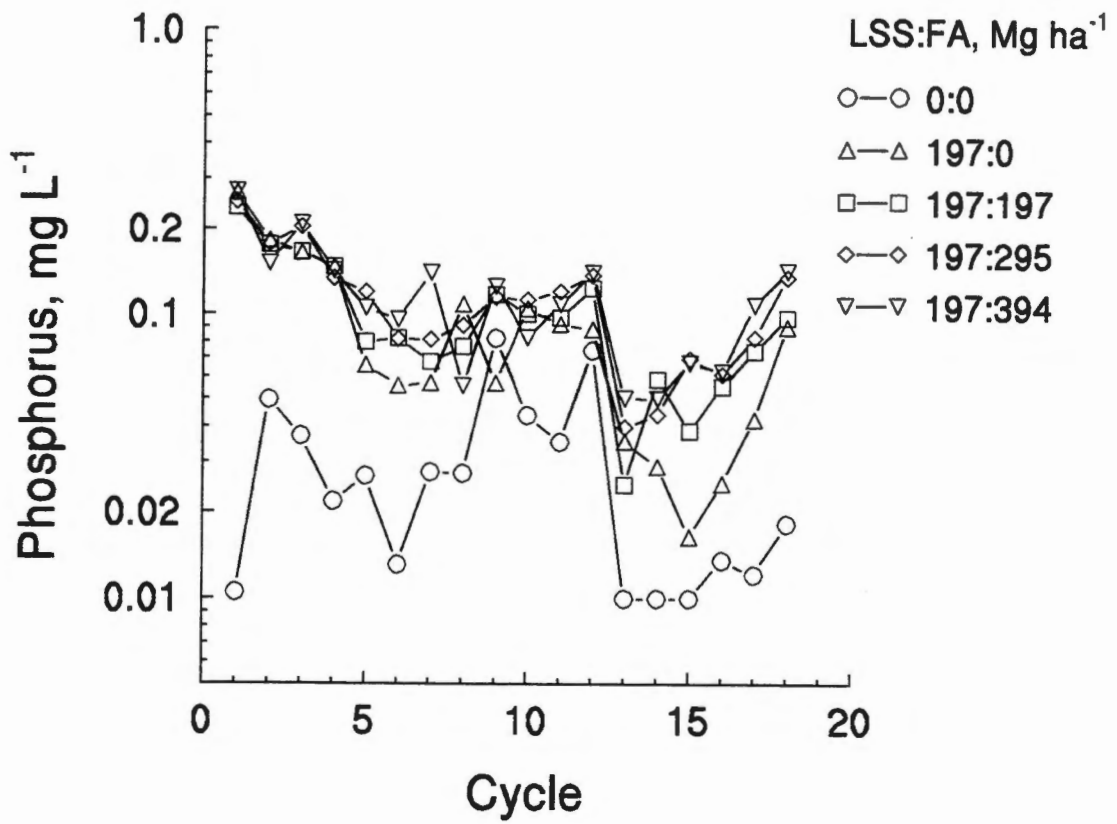


Figure 10. The concentration of phosphorus in humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

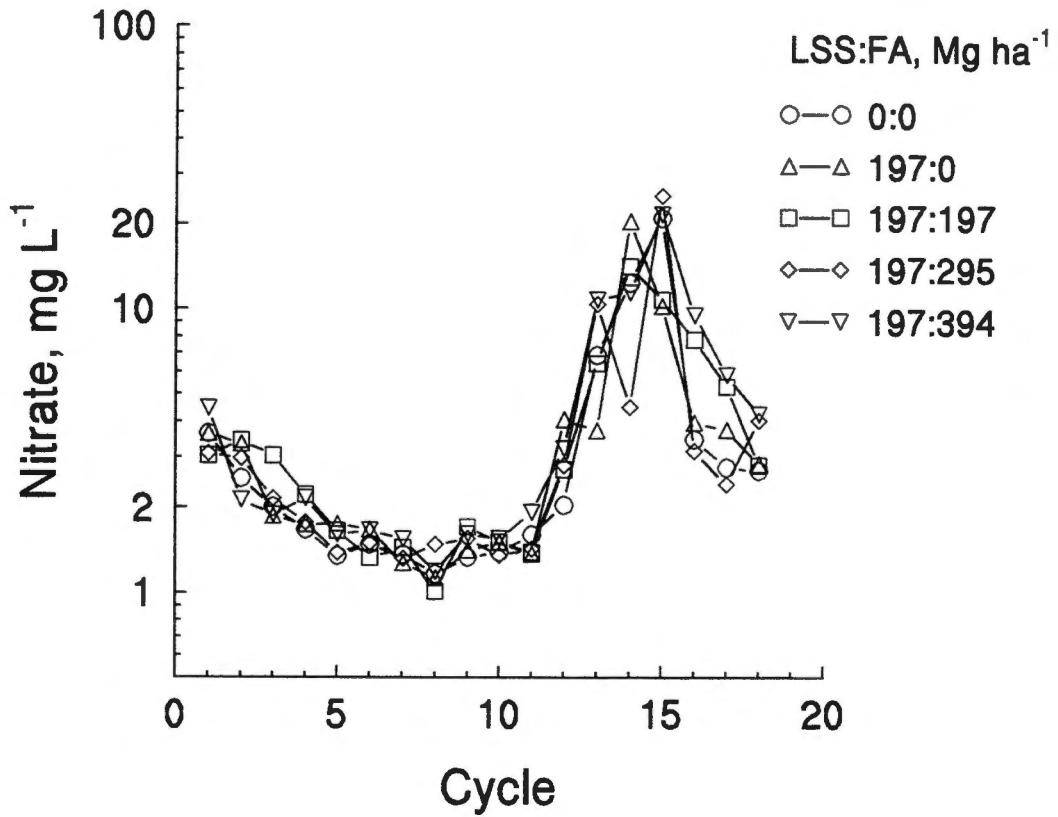


Figure 11. The concentration of nitrate in humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

Potassium concentrations in the unamended mine spoil leachate increased with weathering from an initial concentration of 0.3 mg L^{-1} to approximately 6 mg L^{-1} (Figure 12). Potassium concentrations in the amended spoil leachate initially dropped to approximately 1 mg L^{-1} then oscillated around 2 mg L^{-1} . Acid generation in the unamended mine spoil provided acidic conditions that enhance K-bearing mineral solubilization (e.g. mica). This contributed to the increased K concentrations in the unamended spoil leachates. In general, leachate K concentrations were not influenced by FA rate.

Boron concentrations in the mine spoil leachates increased with increasing fly ash rate (Figure 13). Boron was expected to be found in amended mine spoil leachates because FA particles are generally described as spheroidal borosilicate glass structures. However, during weathering cycles 1 through 4, B concentrations in the LSS amended spoil was elevated relative to the unamended spoil due to the contribution of B from the LSS. The influence of LSS in leachate B concentrations was also observed for the LSS and FA amended mine spoil during the initial weathering cycles.

The difference between amended and unamended mine spoil leachate Al and Fe concentrations was attributed to the LSS applications that increased pH and reduced the solubility of Al- and Fe-bearing minerals. Clearly, Al and Fe

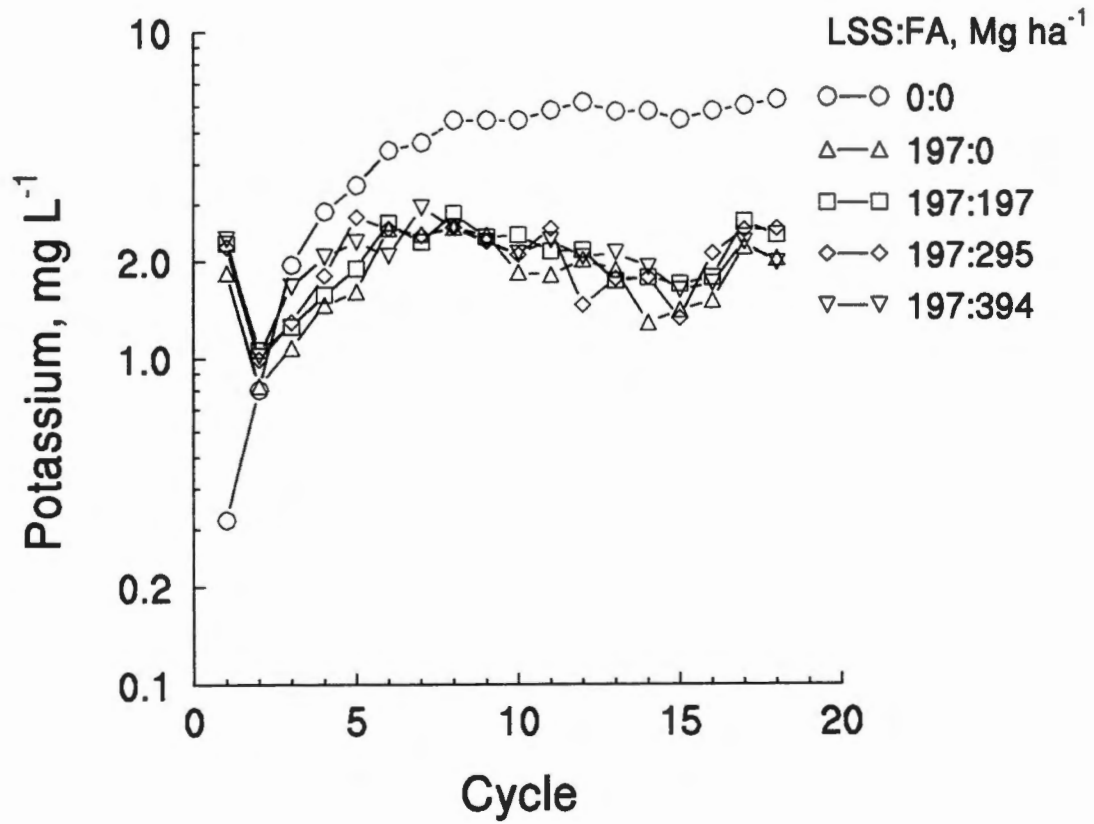


Figure 12. The concentration of potassium in humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

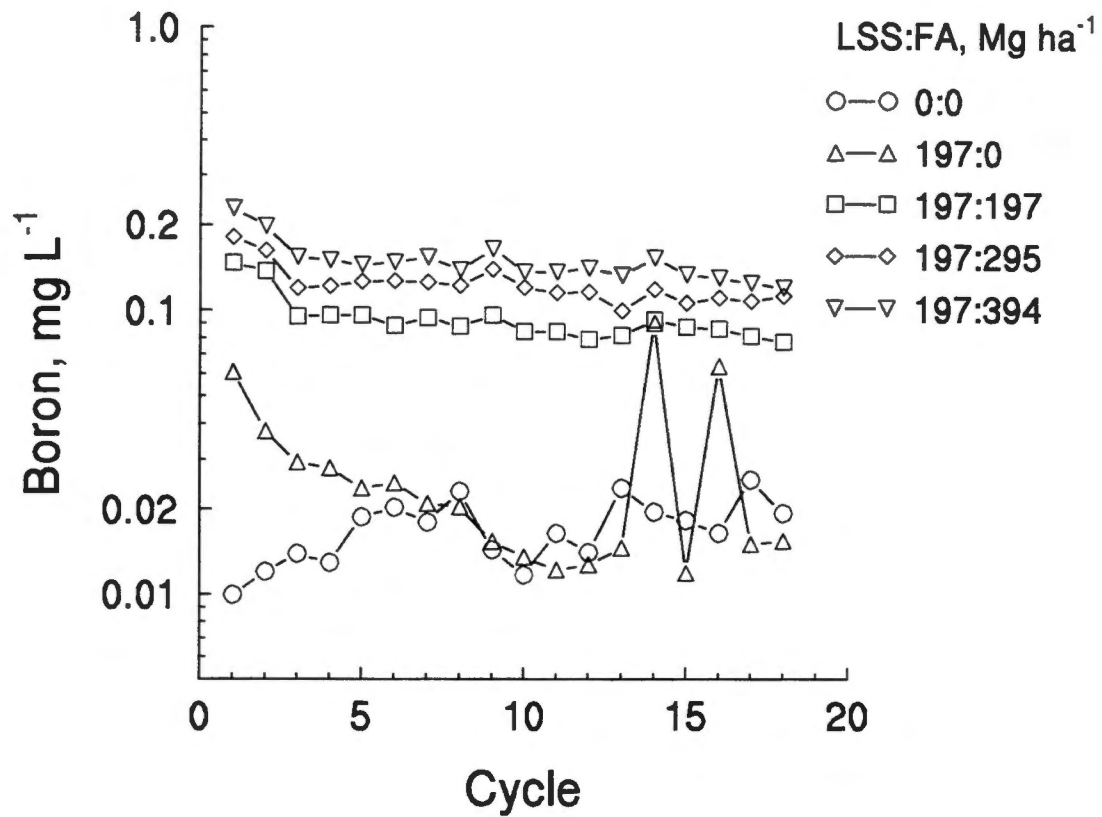


Figure 13. The concentration of boron in humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

concentrations in amended spoil leachates were near the levels of detection after the initial weathering cycle. Aluminum concentrations in the unamended mine spoil leachate decreased from approximately 200 mg L⁻¹ during cycle 1 to approximately 1.0 mg L⁻¹ by cycle 18 (Figure 14). Similarly, Fe concentrations in the unamended mine spoil decreased from 40 mg L⁻¹ to approximately 0.6 mg L⁻¹ by the end of the weathering study (Figure 15). There was no influence of varying FA rate in leachate Al or Fe concentrations.

Like Al and Fe, Mn concentrations were highest in unamended mine spoil leachates. Leachate Mn values in the unamended mine spoil were 30 mg L⁻¹ upon initial weathering, decreasing to approximately 1.0 mg L⁻¹ by the end of the study (Figure 16). However, unlike Al and Fe, leachate Mn decreased gradually in the amended spoil from approximately 20 mg L⁻¹ to levels approaching the detection limit by the end of the weathering study.

The concentrations of Cu were highest in the unamended mine spoil (1.0 mg L⁻¹) during the first weathering cycle due to Cu mobility in low pH systems, but decreased with weathering (Figure 17). Copper concentrations in the amended mine spoil fluctuated about 0.01 mg L⁻¹ for all amended spoil with no apparent influence of FA rate. The difference between amended and unamended spoil Cu leachate

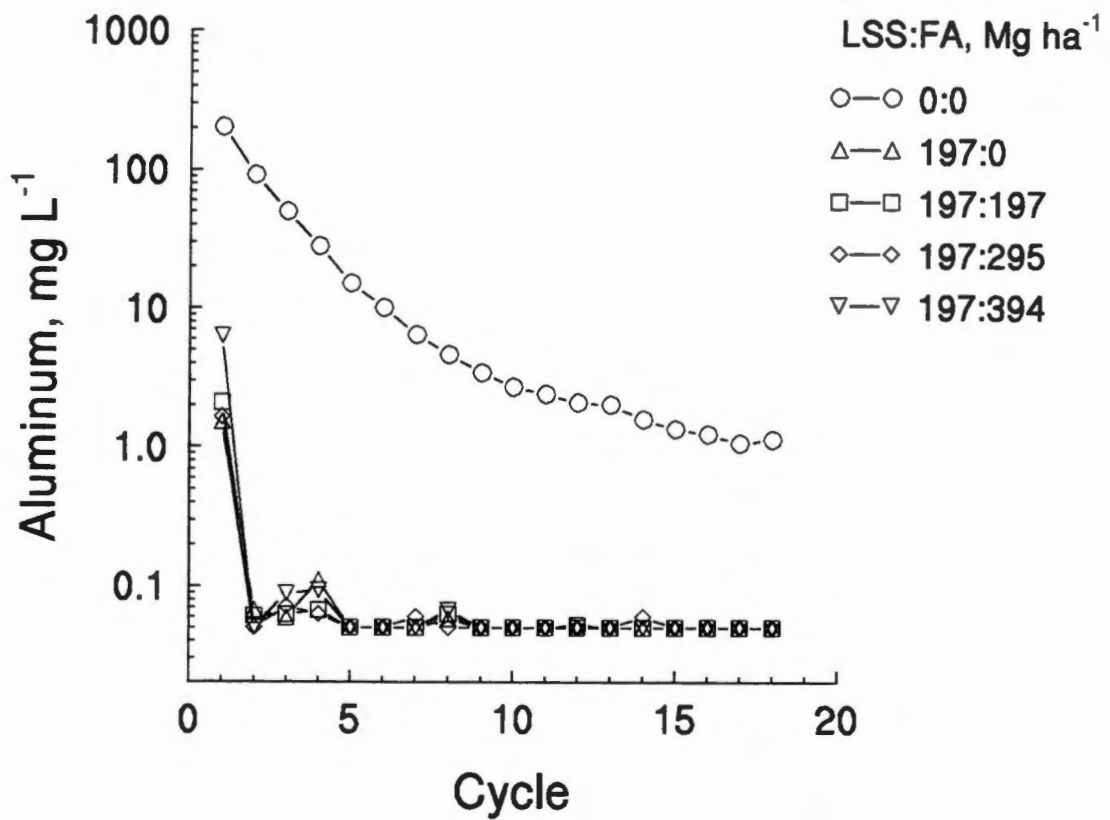


Figure 14. The concentration of aluminum in humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

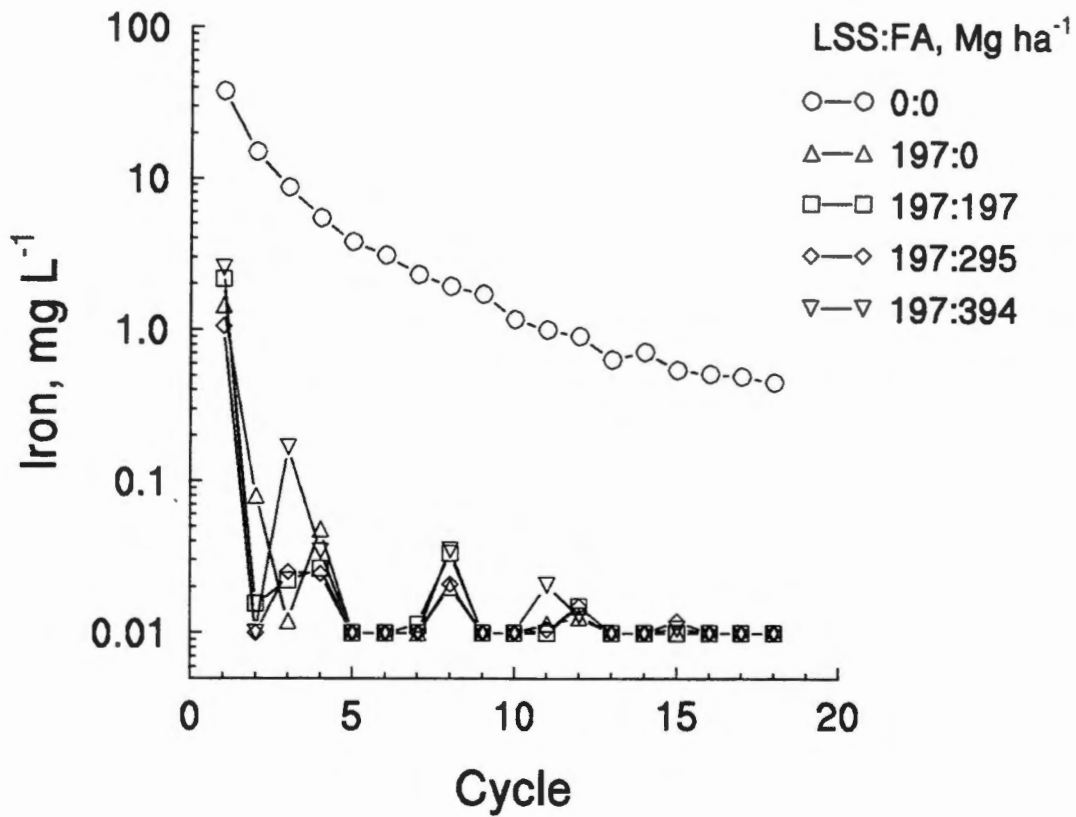


Figure 15. The concentration of iron in humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

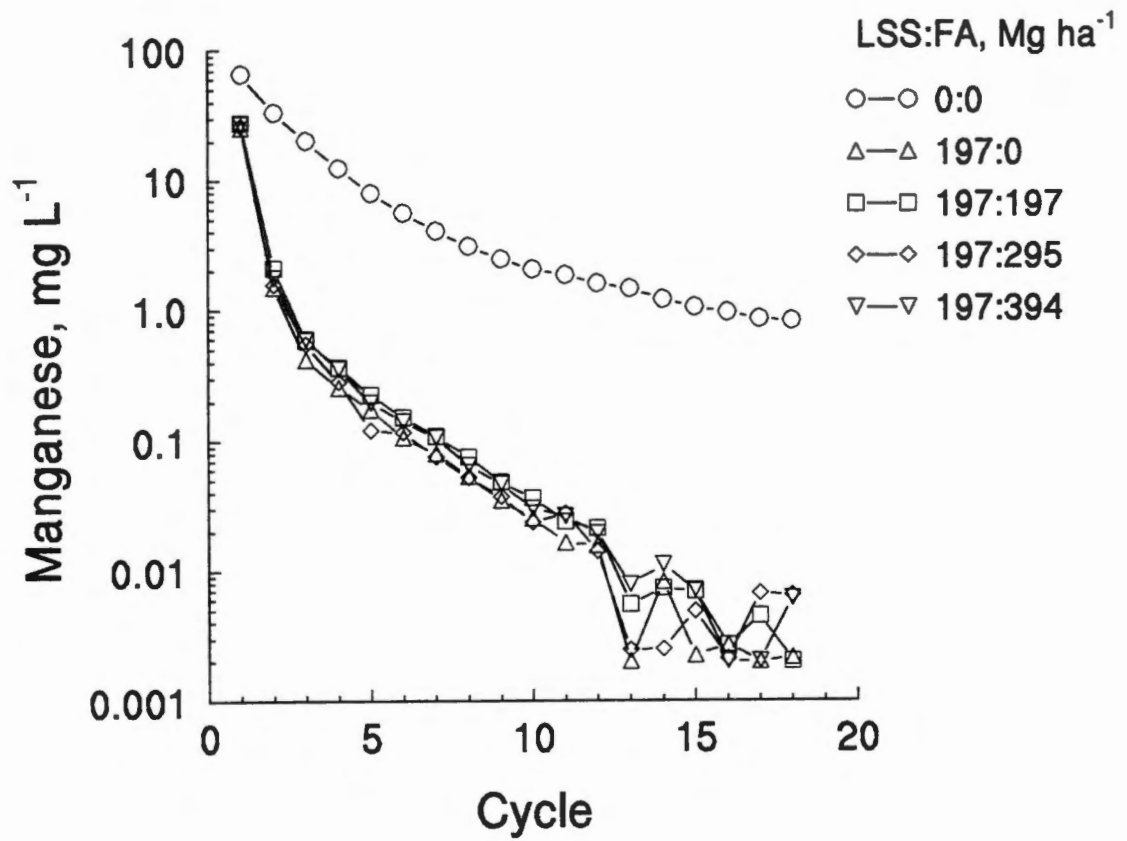


Figure 16. The concentration of manganese in humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

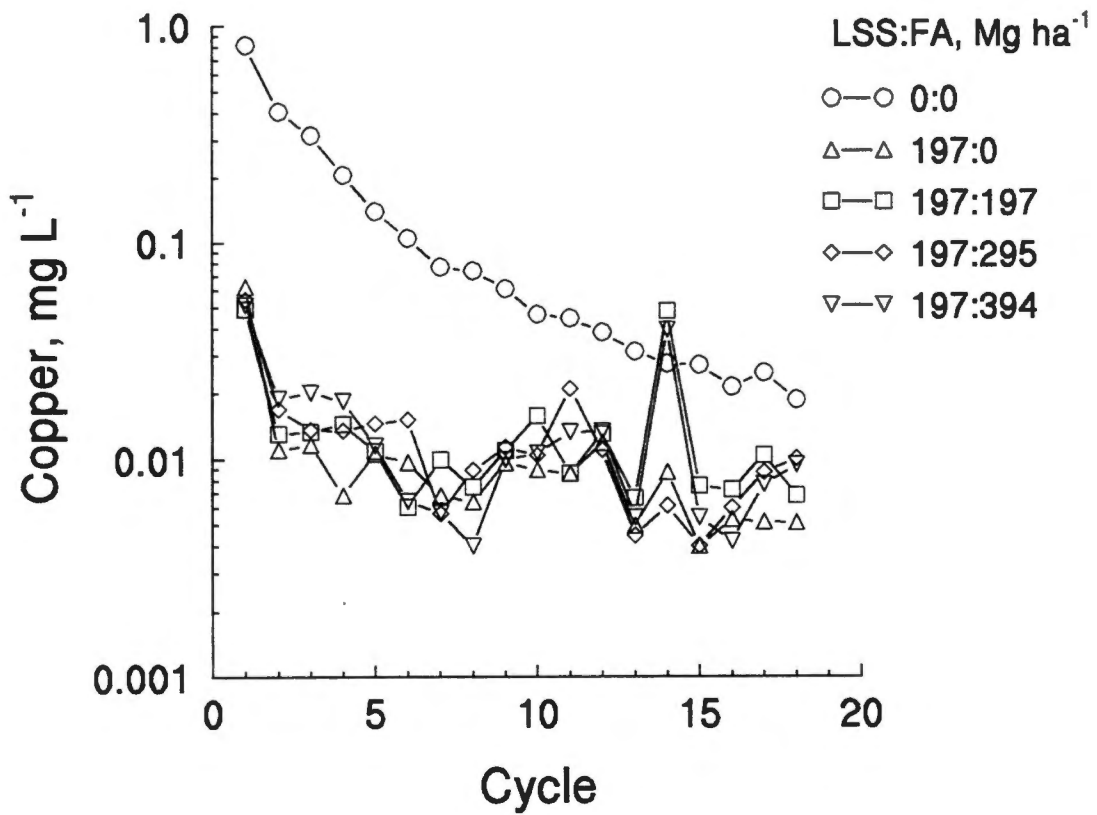


Figure 17. The concentration of copper in humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

concentrations was attributed to the LSS application, increasing pH and the stability of Cu-bearing minerals. Nickel leachate concentrations for the unamended mine spoil were initially found at 3 mg L⁻¹, but decreased to the 0.03 mg L⁻¹ level of detection within 8 weathering cycles (Figure 18). The difference between amended and unamended spoil leachate Ni concentrations was attributed to the LSS application increasing pH and the stability of Ni-bearing minerals. Similar to Ni, Zn in amended spoil leachates were near detectable limits throughout the weathering study, irrespective of FA amendment rate (Figure 19). However, the unamended mine spoil leachate concentrations were initially high in Zn (9.0 mg L⁻¹) and gradually decreased to approximately 0.1 mg L⁻¹ with weathering, as was noted for Cu and Ni, the difference between amended and unamended spoil leachate Zn concentrations was attributed to the LSS application which increased solution pH and the stability of Zn-bearing minerals.

In summary, simulated laboratory weathering showed that the LSS amendment had a dominate impact on leachate chemical characteristics. Leachate pH was effectively neutralized for the duration of the weathering study through the addition of LSS. The decrease of pH observed during the later weathering cycles coincided with fluctuations in

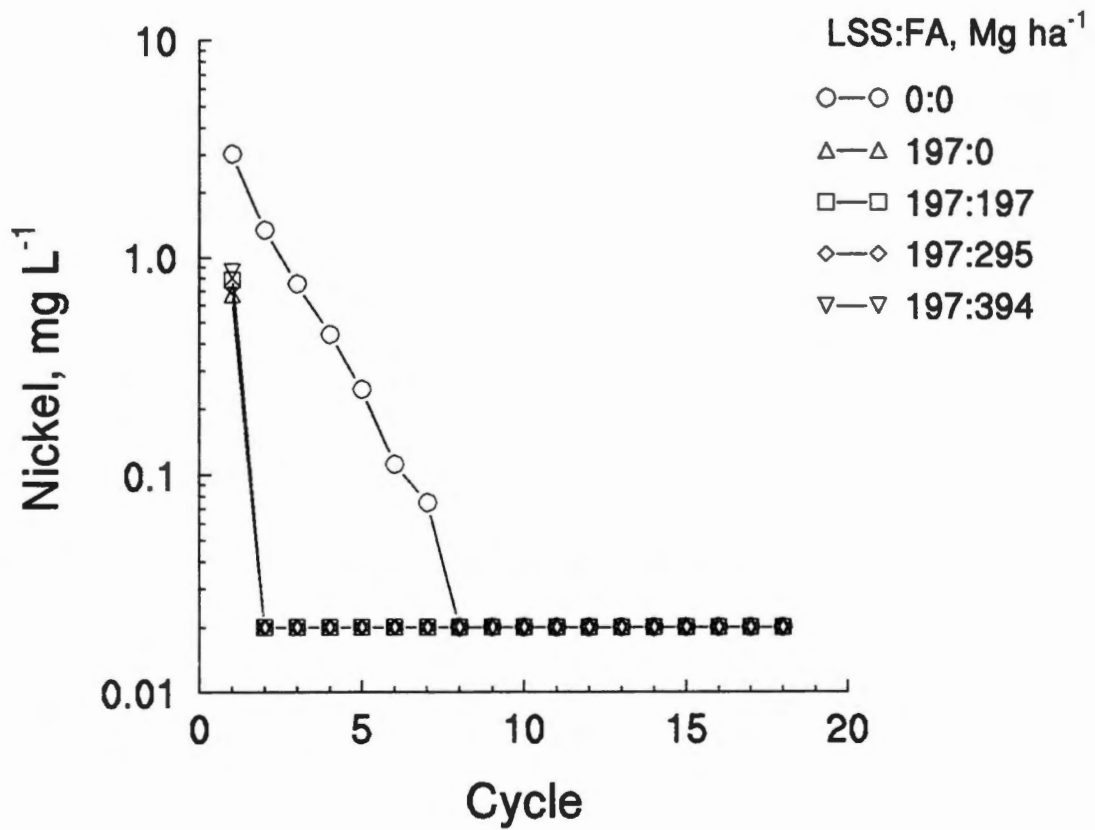


Figure 18. The concentration of nickel in humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

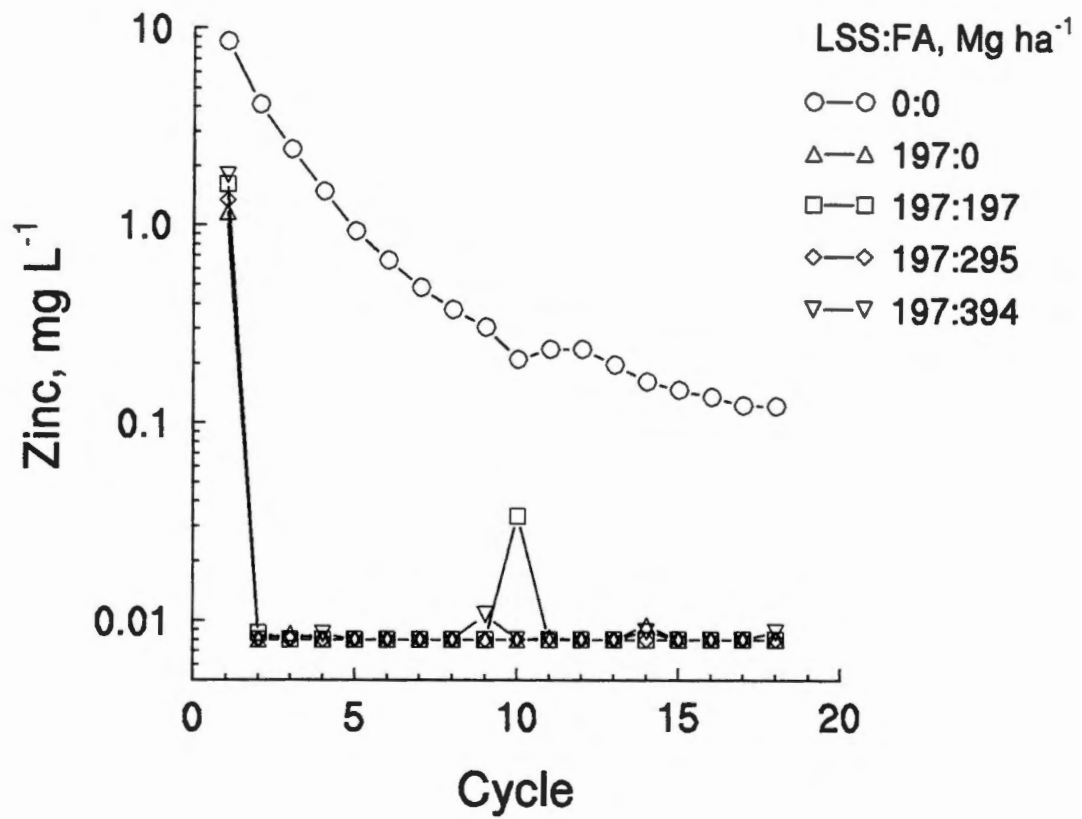


Figure 19. The concentration of zinc in humidity cell leachates from unamended and lime-stabilized sewage sludge (LSS) and fly ash (FA) amended mine spoil.

nitrate concentrations, suggesting increased microbial activity. Sewage sludge was the primary source for calcium, sulfate, chloride, fluoride, and phosphorus in the amended spoil leachates. In general, the leachate concentrations of these elements decreased with weathering. With the exception of phosphorus and boron, FA had no impact on leachate chemistry. Phosphorus and boron were found to increase in the amended mine spoil leachates with increasing FA rates. Conversely, leachate potassium concentrations were greatest in the unamended mine spoil due to dissolution of K-bearing minerals, irrespective of FA rate. Aluminum, iron, manganese, copper, nickel, and zinc concentrations in leachate were highest in the unamended mine spoil due to elevated metal mobility in low pH systems with no apparent influence of FA rates.

Sequential-Selective Dissolution

The sequential selective dissolution (SSD) procedure employed in this study is reported to partition elements into the following operationally-defined chemical pools: soluble-exchangeable (KNO_3 -extractable), adsorbed (H_2O -extractable), organic (NaOH -extractable), carbonate (EDTA -extractable), and sulfide forms (HNO_3 -extractable). The fraction of the total elemental content that is not extracted by the SSD reagents is defined as the residual fraction. The elements subjected to SSD characterization

were Ba, Co, Cr, Cu, Mn, Ni, Pb, Sr, and Zn in the mine spoil and in the LSS/FA-amended mine spoil collected after 2 cycles (unweathered) and after 18 cycles (weathered) of simulated weathering.

For the unweathered material, Ba was found to primarily reside in the residual fraction (Figure 20). Barium in the residual fraction slightly decreased with increasing FA rate with significant increases in the carbonate and soluble-exchangeable forms (Table 2). The sulfide fraction significantly increased upon FA application. The soluble-exchangeable and carbonate forms of Ba were the highest for the highest FA rate. Upon weathering, Ba in the soluble-exchangeable phase was reduced in the FA amended spoil and significantly increased in the carbonate fraction upon LSS application. In general, the residual fraction of Ba was not influenced by weathering. However, in the weathered material, FA rate significantly increased Ba found in the soluble-exchangeable and sulfide forms and decreased in the residual fraction. The application of LSS increased Ba in the adsorbed and carbonate fractions and decreased the organic and residual forms of Ba.

For the unweathered material, Co was primarily found in the residual form which significantly increased upon FA application, irrespective of FA rate (Figure 21). The soluble-exchangeable form of Co was highest in the unamended

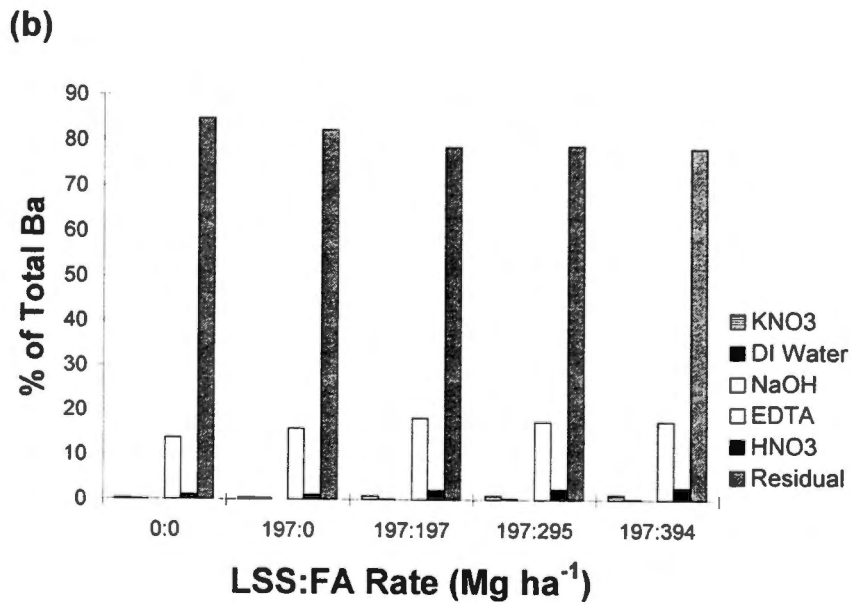
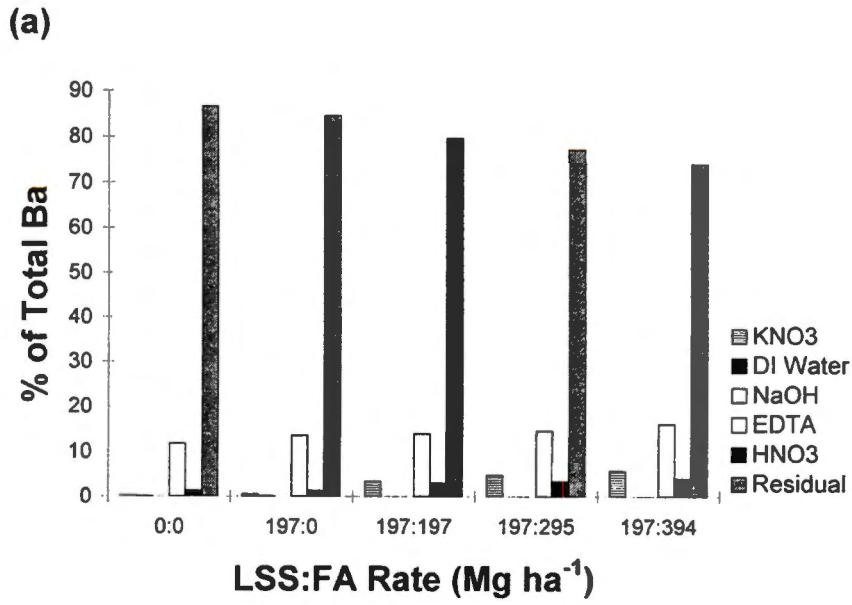


Figure 20. Influence of weathering and lime-stabilized sewage sludge (LSS) and fly ash (FA) amendments on the solid-phase speciation of barium in the mine spoil after (a) 2 cycles and (b) 18 cycles of weathering.

Table 2. Percentages of elements in the extractable fractions of amended mine spoil. Differences between treatments (Trt) and weathering cycles (Cycle) determined by SAS using the general liner model and Student-Newman-Keuls test (95% confidence interval).[†]

Cycle	Trt ^{††}	KNO ₃	KNO ₃ %T [‡]	DI Water	DI %T	NaOH	NaOH %T	EDTA	EDTA %T	HNO ₃	HNO ₃ %T	Residual	Residual %T
----- Barium -----													
2	1	Da	0.20	Ba	0.05	Aa	0.01	Ba	11.82	Ba	1.29	Aa	86.63
	2	Da	0.49	Aa	0.06	Aa	0.01	ABa	13.56	Ba	1.31	Aa	84.57
	3	Ca	3.38	Ba	0.01	Aa	0.01	ABa	13.93	Aa	3.02	Ba	79.66
	4	Ba	4.73	Ba	0.00	Ba	0.00	ABa	14.58	Aa	3.44	Ca	77.24
	5	Aa	5.72	Ba	0.03	Ba	0.00	Aa	16.13	Aa	4.01	Ca	74.10
18	1	Ea	0.25	Ba	0.04	Aa	0.01	Ba	13.82	Ca	1.07	Aa	84.81
	2	Da	0.39	Ab	0.27	Bb	0.00	Aa	15.91	Ca	1.14	Ba	82.30
	3	Cb	0.83	ABb	0.17	Ba	0.00	Ab	18.30	Bb	2.16	Ca	78.54
	4	Bb	0.94	Ab	0.27	Ba	0.00	Ab	17.46	ABb	2.45	Ca	78.87
	5	Ab	1.14	ABb	0.18	Ba	0.00	Aa	17.55	Aa	2.80	Ca	78.33
----- Cobalt -----													
2	1	Aa	13.66	Aa	1.98	Ba	1.21	Aa	6.59	Aa	10.64	Ba	65.91
	2	Aa	11.64	Aa	1.19	Ba	1.25	Aa	7.56	Aa	9.31	Ba	69.05
	3	Ba	6.66	Aa	0.96	ABa	1.92	Aa	6.11	Aa	9.14	Aa	75.21
	4	Ba	5.64	Aa	0.96	ABa	1.79	Aa	6.80	Aa	9.67	Aa	75.14
	5	Ba	4.43	Aa	0.90	Aa	2.35	Aa	6.49	Aa	9.59	Aa	76.25
18	1	Ab	2.77	Ba	1.30	Aa	0.96	Ca	8.86	BCa	11.10	Ab	75.02
	2	Bb	1.12	Aa	1.61	Aa	1.75	Ab	17.40	Ab	14.32	Ba	63.79
	3	Bb	0.87	Ba	1.25	Aa	1.64	Bb	13.73	ABb	12.78	Aa	69.72
	4	Bb	0.77	BCa	1.11	Aa	1.40	BCb	12.04	BCa	11.27	Aa	73.42
	5	Bb	0.59	Ca	0.86	Ab	1.39	BCb	10.99	Ca	8.69	Aa	77.48

[†] For each element, capital letters indicate differences between treatments within the same extractant and the same weathering cycle while lower case letters indicate differences between weathering cycles within the same extractant and the same treatment.

^{††} Treatments: 1, 0:0 Mg ha⁻¹ LSS:FA; 2, 197:0 Mg ha⁻¹ LSS:FA; 3, 197:197 Mg ha⁻¹ LSS:FA;

4, 197:295 Mg ha⁻¹ LSS:FA; 5, 197:394 Mg ha⁻¹ LSS:FA.

[‡] Extracted/Total*100 = %T

Table 2. (continued)

Cycle	Trt	KNO ₃	KNO ₃ %T	DI Water	DI %T	NaOH	NaOH %T	EDTA	EDTA %T	HNO ₃	HNO ₃ %T	Residual	Residual %T
Chromium													
2	1	Aa	0.36	Aa	0.53	Aa	0.39	Aa	3.81	Aa	4.51	Ba	90.40
	2	Aa	0.38	Aa	0.55	Aa	0.39	Ba	1.18	Aa	4.81	Aa	92.69
	3	Aa	0.35	Aa	0.50	Aa	0.34	Ba	0.85	Aa	4.62	Aa	93.35
	4	Aa	0.34	Aa	0.50	Aa	0.33	Ba	1.16	Aa	4.91	Aa	92.77
	5	Aa	0.35	Aa	0.50	Aa	0.34	Ba	1.14	Aa	5.17	Aa	92.50
18	1	Aa	0.37	Ba	0.52	Aa	0.35	Ab	6.34	Da	5.17	Cb	87.25
	2	Aa	0.37	Ab	1.18	Ba	0.31	Ba	1.08	ABb	6.20	ABb	90.86
	3	Aa	0.37	Ab	1.11	Ba	0.31	Ba	1.41	Ab	6.47	Bb	90.34
	4	Aa	0.44	Ab	1.15	BCa	0.29	Ba	1.21	BCa	5.77	ABb	91.14
	5	Aa	0.32	Ab	0.96	Ca	0.27	Ba	1.14	CDa	5.33	Ab	91.98
Copper													
2	1	Aa	68.74	Aa	0.20	Ca	11.10	Ca	14.09	Ba	5.88	Ba	0.00
	2	Ba	18.43	Aa	0.14	Aa	36.56	Aa	37.29	ABa	7.58	Ba	0.00
	3	Ca	12.81	Aa	0.03	Ba	24.26	Ba	25.42	ABa	8.56	Aa	28.92
	4	Ca	7.69	Aa	0.01	Ba	22.41	Ba	24.95	ABa	9.13	Aa	35.80
	5	Ca	5.94	Aa	0.09	Ba	20.82	Ba	23.06	Aa	10.14	Aa	39.95
18	1	Ab	21.04	Ba	0.00	Aa	13.57	Ab	21.43	Ab	10.86	Bb	33.09
	2	Bb	0.23	Ab	0.44	Ab	13.04	Ab	25.54	Ba	5.55	Ab	55.19
	3	Bb	0.21	Ab	0.55	Ab	13.45	Aa	26.84	Ba	6.99	Aa	51.96
	4	Bb	0.20	Ab	0.51	Ab	12.60	Aa	24.83	Ba	7.15	Ab	54.71
	5	Ba	0.15	Aa	0.40	Ab	10.66	Aa	20.69	Bb	6.09	Ab	62.00

Table 2. (continued)

Cycle	Trt	KNO ₃ %T	DI Water	DI %T	NaOH	NaOH %T	EDTA	EDTA %T	HNO ₃	HNO ₃ %T	Residual	Residual %T
..... Lead												
2	1	Aa	Aa	1.40	Ba	0.88	Aa	20.08	Ba	2.72	Aa	72.25
	2	Ba	Aa	1.45	Ba	1.00	Aa	20.87	Ba	2.50	Aa	73.17
	3	Ba	Aa	1.35	Aa	1.44	Aa	18.91	ABa	3.56	Aa	73.79
	4	Ba	Aa	1.30	Ba	1.13	Aa	18.92	Aa	3.84	Aa	73.90
	5	Ba	Aa	1.22	Aa	1.62	Aa	19.80	Aa	4.25	Aa	72.25
18	1	Ab	Ba	1.08	Aa	0.78	ABa	19.10	Aa	3.08	ABa	73.93
	2	Ba	Aa	1.39	Aa	1.18	Aa	21.89	Aa	3.66	Ba	70.91
	3	Ba	ABa	1.19	Aa	1.14	ABa	18.64	Aa	3.73	ABa	74.47
	4	Ba	ABa	1.19	Aa	1.03	ABa	18.53	Aa	3.80	ABa	74.62
	5	Ba	Ba	1.00	Ab	1.03	Ba	16.34	Aa	3.22	Aa	77.71
..... Manganese												
2	1	Aa	Aa	0.00	Aa	0.09	Aa	6.70	Aa	9.24	Aa	34.05
	2	ABa	Aa	0.00	Aa	0.09	Aa	7.28	Aa	7.00	Aa	48.26
	3	ABa	Aa	0.00	Aa	0.15	Aa	9.41	Aa	9.96	Aa	32.95
	4	ABa	Aa	0.00	Aa	0.16	Aa	8.78	Aa	8.55	Aa	45.73
	5	Ba	Aa	0.00	Aa	0.09	Aa	7.92	Aa	7.43	Aa	56.65
18	1	Ab	Ba	0.00	Aa	0.04	Ba	6.35	Aa	4.19	Ab	84.69
	2	Bb	Ab	0.23	Aa	0.06	Ab	56.12	Aa	9.43	Ba	33.75
	3	Bb	Ab	0.30	Aa	0.06	Ab	69.37	Aa	11.80	Ba	17.90
	4	Bb	Ab	0.26	Ab	0.04	Ab	65.05	Aa	11.46	Ba	22.43
	5	Bb	Ab	0.19	Aa	0.07	Ab	50.37	Aa	8.69	Ba	40.22

Table 2. (continued)

Cycle	Trt	KNO ₃	KNO ₃ %T	DI Water	DI %T	NaOH	NaOH %T	EDTA	EDTA %T	HNO ₃	HNO ₃ %T	Residual	Residual %T
..... Nickel													
2	1	Aa	21.82	Aa	3.53	Aa	3.16	Aa	17.94	Aa	20.42	Ba	33.13
	2	Aa	33.57	Aa	4.76	Aa	4.98	Aa	31.17	Aa	25.52	Ba	0.00
	3	Ba	11.60	Ba	2.48	Aa	3.71	Ba	17.21	Ba	14.85	Aa	50.15
	4	CBa	8.57	Ba	2.21	Ba	2.60	Ba	16.53	Ba	14.02	Aa	56.07
	5	Ca	7.32	Ba	2.16	Ba	3.25	Ba	16.90	Ba	14.67	Aa	55.70
18	1	Ab	6.55	Ab	6.27	Ca	4.32	Da	28.98	Ab	35.79	Bb	18.08
	2	Bb	3.50	Ba	5.02	Ab	7.90	Ab	53.79	Ba	29.79	Ca	0.00
	3	Bb	2.82	Cb	4.09	Bb	6.38	Bb	45.18	Bb	26.46	Bb	15.07
	4	Bb	2.17	Da	3.13	Cb	4.83	Cb	36.41	Ca	19.75	Ab	33.71
	5	Bb	1.70	Ea	2.46	Ca	3.90	CDb	30.65	Ca	16.10	Ab	45.20
..... Strontium													
2	1	Da	0.08	Aa	0.15	Aa	0.04	Aa	25.88	Da	0.37	Aa	73.49
	2	Ca	1.89	Aa	0.01	Aa	0.04	Ba	21.73	Da	0.26	Aa	76.07
	3	Ba	4.40	Aa	0.00	Ba	0.02	Ba	20.17	Ca	4.84	Aa	70.57
	4	ABa	4.82	Aa	0.00	Ba	0.02	Ba	19.06	Ba	5.42	Aa	70.67
	5	Aa	5.11	Aa	0.00	Ba	0.02	Ba	18.38	Aa	6.09	Aa	70.40
18	1	Ca	0.08	Ab	0.03	ABb	0.03	Ab	16.23	Ba	0.38	Ab	83.25
	2	Bb	0.95	Ba	0.00	Aa	0.03	Aa	17.01	Ba	0.50	Aa	81.50
	3	Ab	2.87	Ba	0.00	BCa	0.02	Aa	19.14	Ab	3.08	Aa	74.89
	4	Ab	3.12	Ba	0.00	BCa	0.03	Aa	16.58	Ab	3.64	Aa	76.64
	5	Ab	2.69	Ba	0.00	Ca	0.01	Aa	13.06	Ab	3.18	Aa	81.05

Table 2. (continued)

Cycle	Trt	KNO ₃	KNO ₃ %T	DI Water	DI %T	NaOH	NaOH %T	EDTA	EDTA %T	HNO ₃	HNO ₃ %T	Residual	Residual %T	
								Zinc						
2	1	Aa	17.17	Aa	0.09	Aa	0.38	Ca	10.12	Aa	14.39	Aa	57.85	
	2	Ba	16.78	Aa	0.00	Aa	0.41	Ba	12.04	Aa	14.19	Aa	56.58	
	3	ABa	14.99	Aa	0.01	Aa	0.48	Aa	14.44	Aa	14.01	Aa	56.06	
	4	Ba	12.73	Aa	0.00	Aa	0.48	Aa	15.25	Aa	12.92	Aa	58.62	
	5	Ba	11.13	Aa	0.28	Aa	0.57	Aa	16.09	Aa	13.15	Aa	58.78	
18	1	Ab	5.72	Ca	0.03	Aa	0.41	Ba	10.49	Ab	17.60	Aa	65.75	
	2	Bb	0.19	Ab	0.69	Aa	0.36	Ab	28.51	Ab	17.82	Ba	52.43	
	3	Bb	0.19	Ab	0.67	Aa	0.57	Ab	31.60	Ab	18.08	Ba	48.88	
	4	Bb	0.27	ABb	0.59	Aa	0.65	Ab	27.82	ABb	14.89	Ba	55.77	
	5	Bb	0.16	Ba	0.49	Aa	0.63	Ab	26.77	Ba	12.72	Ba	59.24	

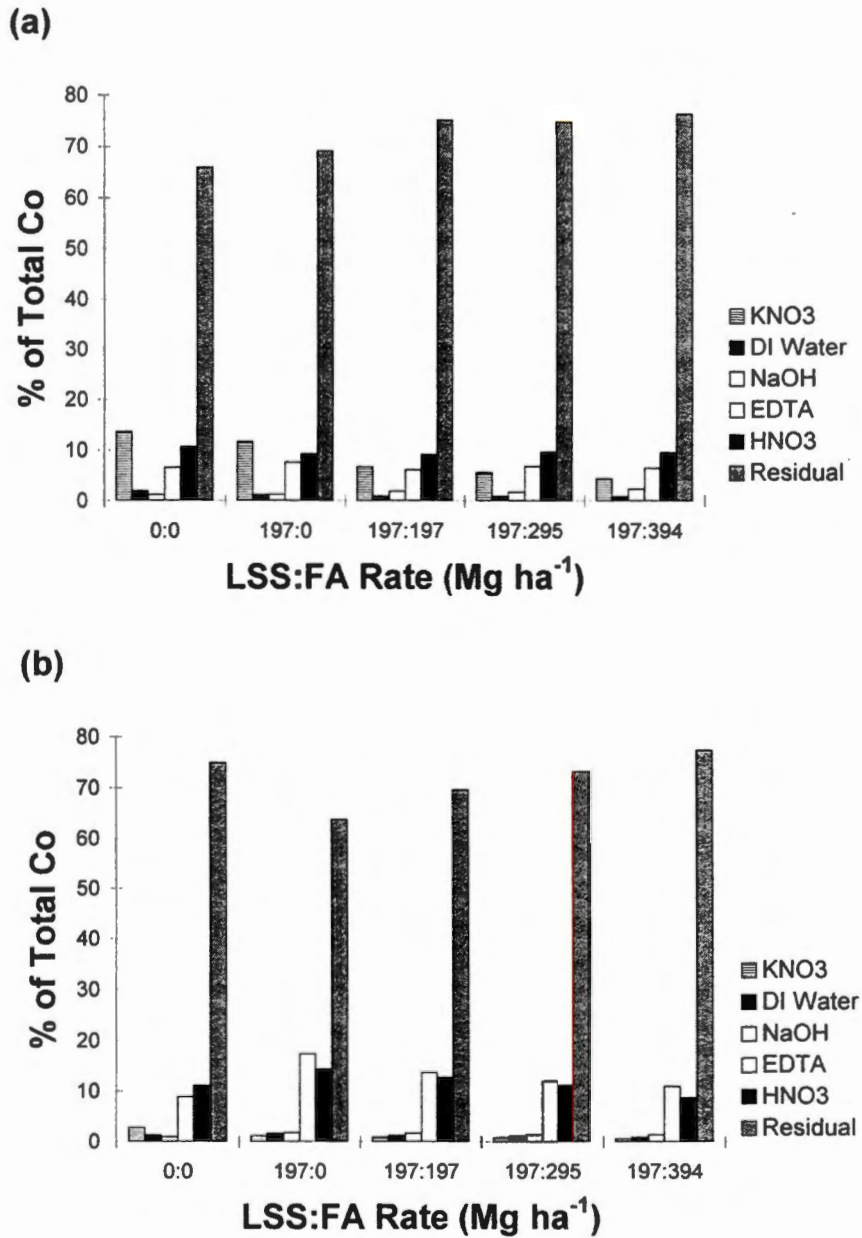


Figure 21. Influence of weathering and lime-stabilized sewage sludge (LSS) and fly ash (FA) amendments on the solid-phase speciation of cobalt in the mine spoil after (a) 2 cycles and (b) 18 cycles of weathering.

spoil (13.6%) and significantly decreased upon FA application (Table 2). Weathering significantly shifted Co speciation from the soluble-exchangeable form to the carbonate form. However, the soluble-exchangeable form of Co in the unamended mine spoil was significantly higher than that in the amended mine spoil, irrespective of weathering. In the weathered and amended mine spoil, the soluble-exchangeable form of Co significantly decreased upon LSS application, irrespective of FA rate.

For the unweathered material, Cr was primarily found in the residual phase (approximately 90%) and the sulfide form (between 5 and 11%) for unamended and amended mine spoil (Figure 22). The unamended mine spoil showed approximately equal proportions of Cr in the carbonate and sulfide fractions (5%). For the unweathered mine spoil, LSS application significantly decreased the carbonate form of Cr and significantly increased the residual forms (Table 2). In general, weathering significantly decreased the residual fraction and significantly increased in the adsorbed fraction of Cr. For the unweathered and weathered mine spoil, LSS application shifted Cr from the carbonate fraction to the sulfide and residual fractions.

For the unweathered and unamended mine spoil, Cu was primarily found in the soluble-exchangeable form at approximately 69% (Figure 23). A significant decrease in

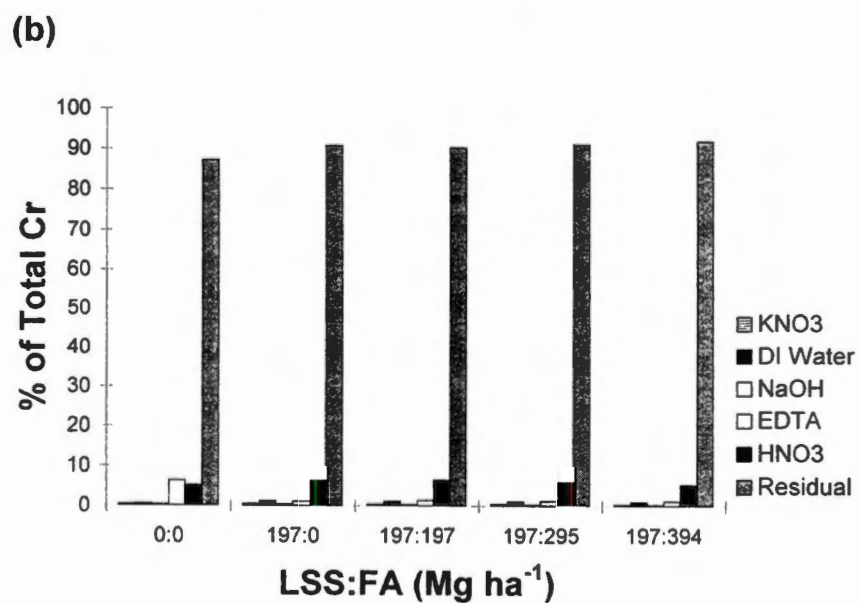
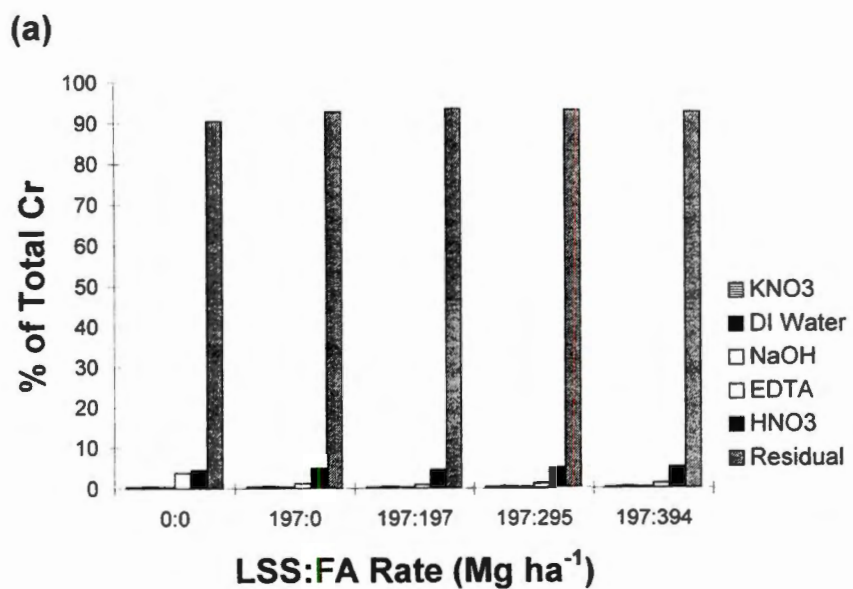


Figure 22. Influence of weathering and lime-stabilized sewage sludge (LSS) and fly ash (FA) amendments on the solid-phase speciation of chromium in the mine spoil after (a) 2 cycles and (b) 18 cycles of weathering.

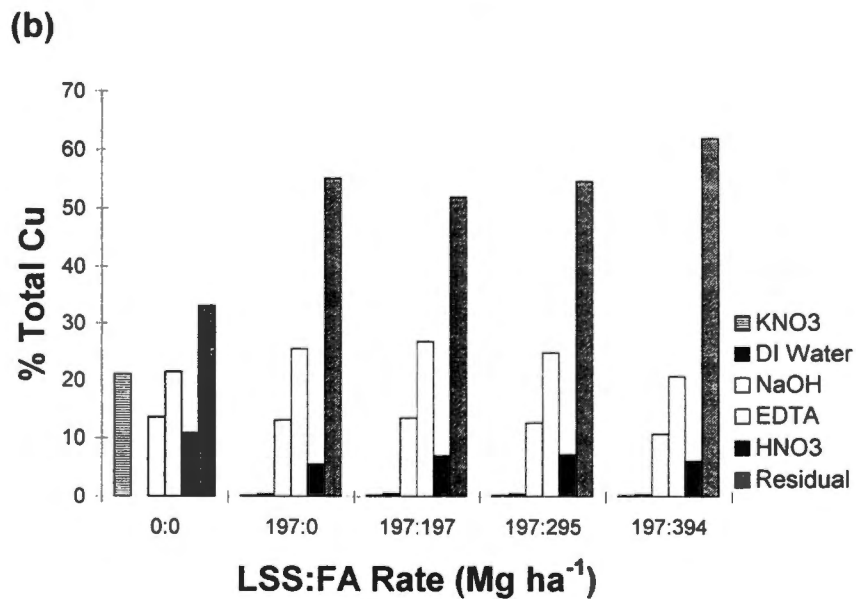
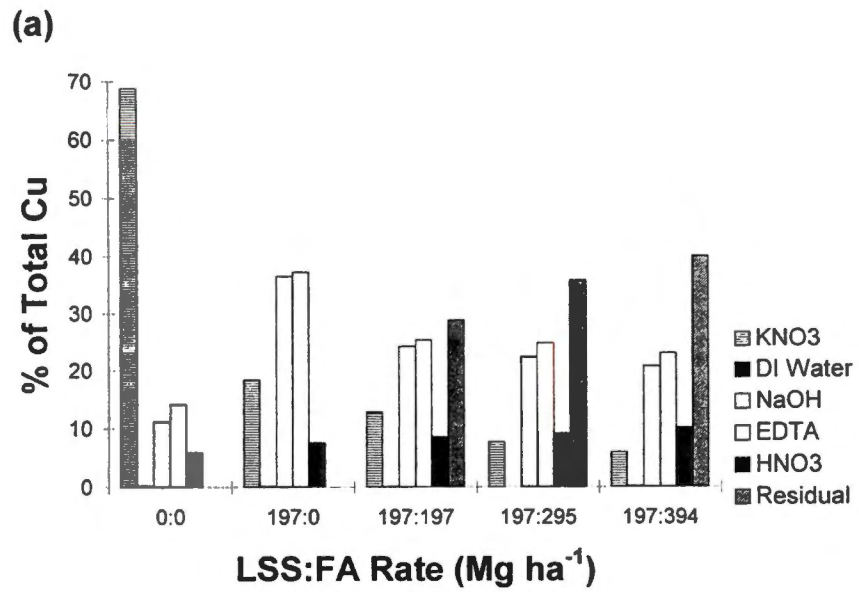


Figure 23. Influence of weathering and lime-stabilized sewage sludge (LSS) and fly ash (FA) amendments on the solid-phase speciation of copper in the mine spoil after (a) 2 cycles and (b) 18 cycles of weathering.

the soluble-exchangeable form was observed upon LSS application and FA rate (Table 2). The organic, carbonate, and sulfide forms of Cu were also present in the unamended spoil (11%, 14%, and 6%, respectively). The LSS amended mine spoil showed that Cu primarily resided in the organic (37 %) and carbonate forms (37%) with 18% of the extracted Cu residing in the soluble-exchangeable fraction. In general, increasing FA rates significantly decreased Cu in the soluble-exchangeable form and impacted the distribution of Cu in the organic, carbonate, and residual forms. In general, weathering allowed all Cu forms to shift to the residual forms. More specifically, weathering allowed the residual and adsorbed fractions of Cu to increase significantly upon LSS application, irrespective of FA rate. Further, LSS application significantly decreased the fraction of Cu found in the soluble-exchangeable and sulfide forms with weathering. Fly ash application rate did not influence the distribution of Cu in the weathered materials.

For the unweathered material, Mn was predominately found in the soluble-exchangeable and residual fractions, irrespective of amendment (Figure 24 and Table 2). Upon weathering, the unamended spoil decreased Mn found in the soluble-exchangeable and carbonate forms with a significant shift to the residual fraction. In all LSS-amended mine spoil, Mn shifted from the residual fraction to the carbonate form upon weathering, irrespective of FA rate.

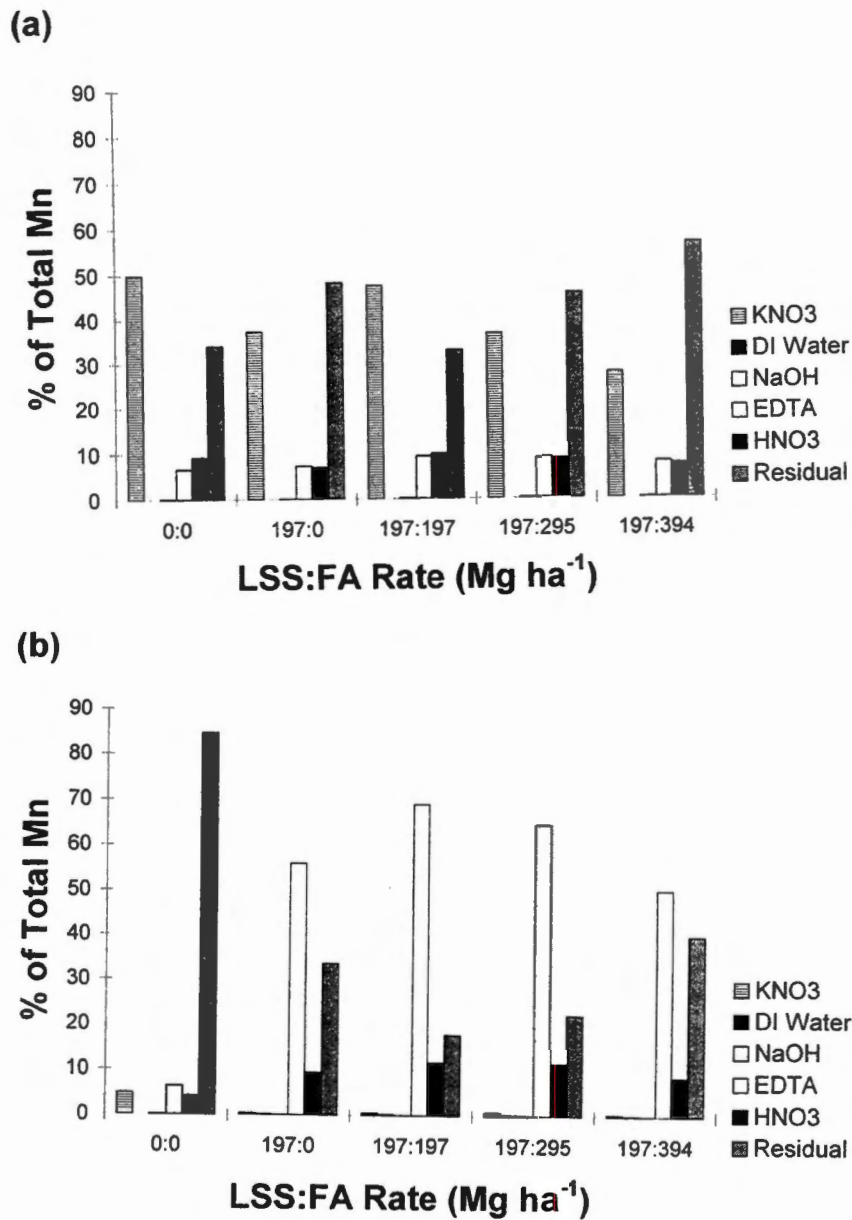


Figure 24. Influence of weathering and lime-stabilized sewage sludge (LSS) and fly ash (FA) amendments on the solid-phase speciation of manganese in the mine spoil after (a) 2 cycles and (b) 18 cycles of weathering.

For the unweathered and unamended spoil, Ni was highest in the residual fraction (33%), with approximately equal proportions in the adsorbed and organic fractions (approximately 3%) (Figure 25). Further, the soluble-extractable form of Ni was found to be 22%, with the carbonate and sulfide fractions containing 18% and 20%, respectively. For the unweathered and LSS-amended spoil, Ni was found in the soluble-exchangeable (34%), carbonate (31%), and sulfide forms (26%). Further, the adsorbed and organic forms of Ni were approximately in equal proportions (5%). In general, increasing FA rate increased Ni found in the residual form and decreased Ni found in the adsorbed, sulfide, and carbonate forms (Table 2). Upon weathering, Ni in the soluble-exchangeable and residual fractions decreased with subsequent increases in the carbonate and sulfide fractions. For the weathered and LSS-amended spoil, the soluble-exchangeable form of Ni decreased with subsequent increases in the carbonate and sulfide forms. In general, the residual forms of Ni decreased with weathering for the FA-amended mine spoil while increasing in the carbonate and sulfide fractions. In general, carbonate forms of Ni decreased with increasing FA rate. Further, sulfide forms of Ni increased upon weathering, yet decreased fairly linearly with FA rate.

Lead was primarily found in the residual fraction,

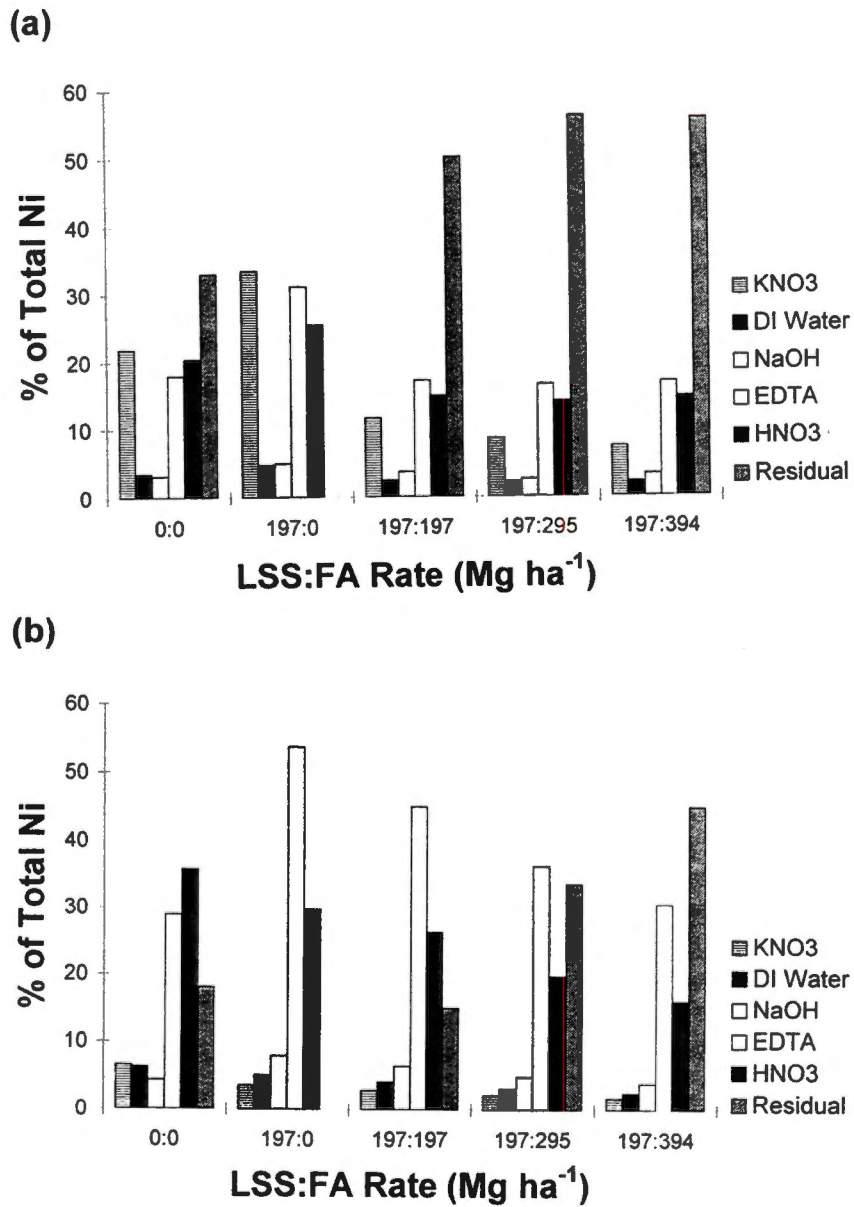


Figure 25. Influence of weathering and lime-stabilized sewage sludge (LSS) and fly ash (FA) amendments on the solid-phase speciation of nickel in the mine spoil after (a) 2 cycles and (b) 18 cycles of weathering.

irrespective of FA rate (Figure 26 and Table 2). The soluble-exchangeable form of Pb was found to decrease upon LSS application, irrespective of FA rate. Weathering did not impact Pb distribution in the amended mine spoil. However, the soluble-exchangeable forms of Pb were observed to decrease upon weathering and LSS amendment. In general, weathering had no influence on Pb distribution, irrespective of amendment.

Similar to Pb, Sr was primarily found in the residual fraction of the unweathered mine spoil, ranging from 70% to 76% (Figure 27). Strontium in the carbonate fraction ranged from 18% to 26%. In general, as FA rate increased, Sr was found to increase in the soluble-exchangeable and sulfide fractions (Table 2). Further, upon FA application, Sr was found to decrease in the organic form. However, LSS application significantly decreased the amount of Sr found in the carbonate fraction. Weathering resulted in significant decrease in the soluble-exchangeable form of Sr. Although the residual and the carbonate fractions of Sr predominated, FA application tended to increase Sr found in the soluble-exchangeable forms. In general, for FA amended mine spoil, the soluble-exchangeable forms of Sr shifted with weathering to the sulfide form.

For the unweathered spoil material, Zn was predominately found the residual form (58% on average). The

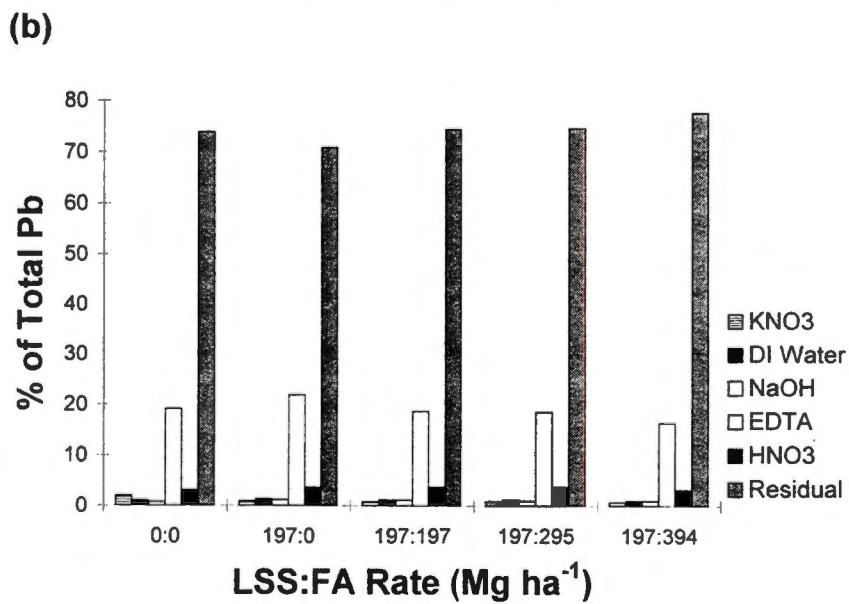
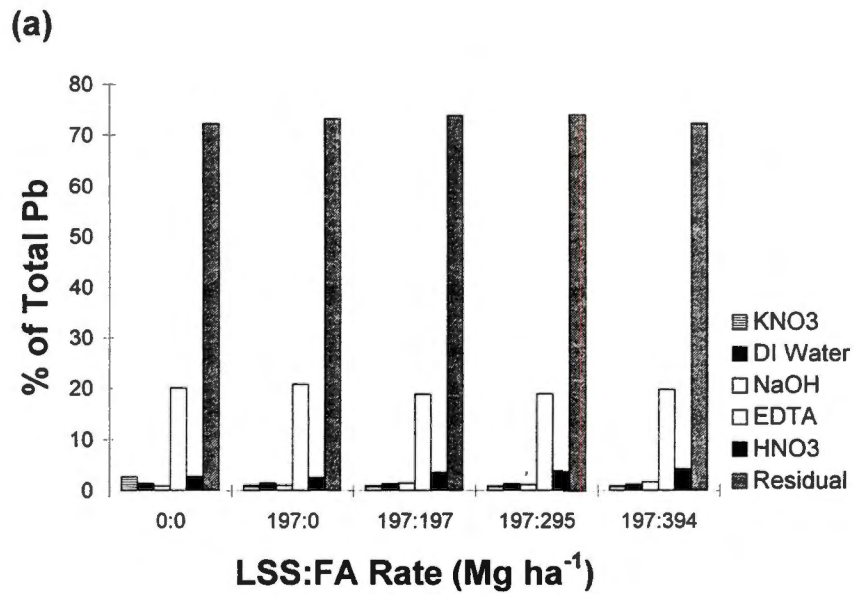


Figure 26. Influence of weathering and lime-stabilized sewage sludge (LSS) and fly ash (FA) amendments on the solid-phase speciation of lead in the mine spoil after (a) 2 cycles and (b) 18 cycles of weathering.

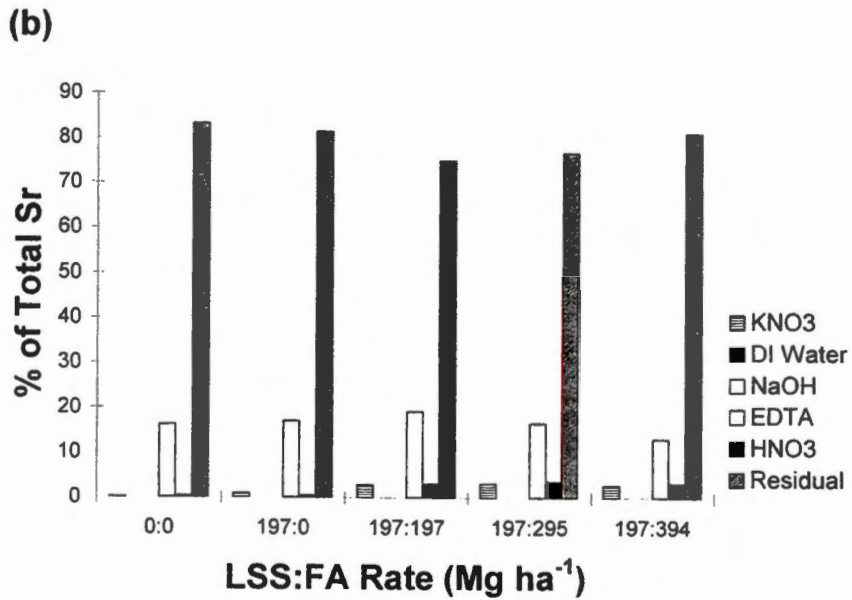
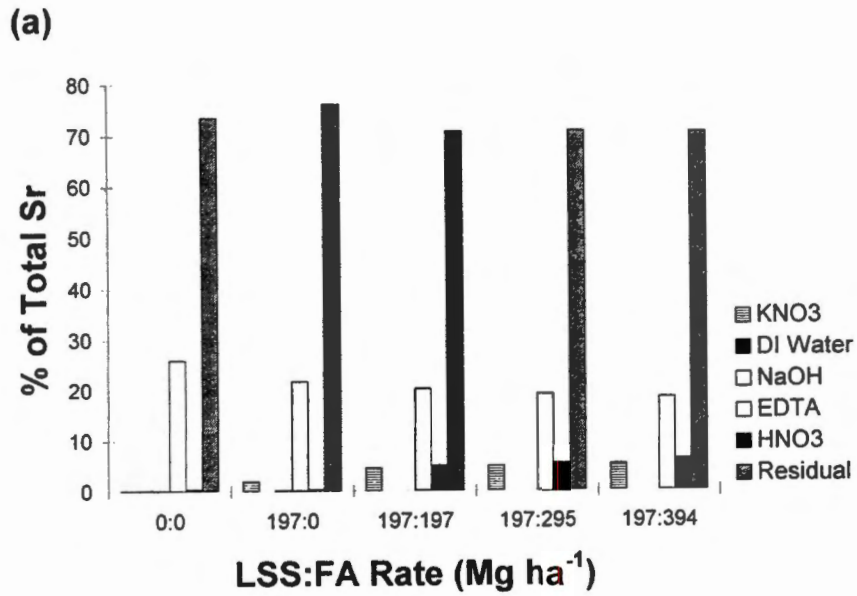


Figure 27. Influence of weathering and lime-stabilized sewage sludge (LSS) and fly ash (FA) amendments on the solid-phase speciation of strontium in the mine spoil after (a) 2 cycles and (b) 18 cycles of weathering.

soluble-exchangeable and residual forms of Zn were influenced by amendment. However, a significant increase in Zn found in the carbonate form was observed with FA amendment (Figure 28). Upon weathering, Zn in all LSS-amended mine spoil increased in the carbonate fraction with a corresponding decrease in the soluble-exchangeable and sulfide forms, generally irrespective of FA rate.

In summary, the SSD procedure used in this study examined the solid-phase speciation of Ba, Co, Cr, Cu, Mn, Ni, Pb, Sr, and Zn. The speciation of Ba increased in the soluble-exchangeable and carbonate fractions with increasing FA rate. Strontium increased in the soluble-exchangeable and sulfide fractions with increasing FA rate. Weathering tended to shift Ba to the carbonate fraction and Sr showed no significant shifts to any other phase. Chromium found in the residual fraction significantly increased upon LSS application. Weathering of the LSS-amended mine spoil increased Cr found in the adsorbed fraction, yet maintained predominance in the residual fraction. Chromium solid-phase speciation was not impacted by FA application or weathering. Although Co was primarily found in the residual fraction, Co showed a significant decrease in the soluble-exchangeable form upon LSS and FA application in the weathered material. In general, FA did not impact Co solid-phase speciation. However, LSS application decreased Co found in the soluble-

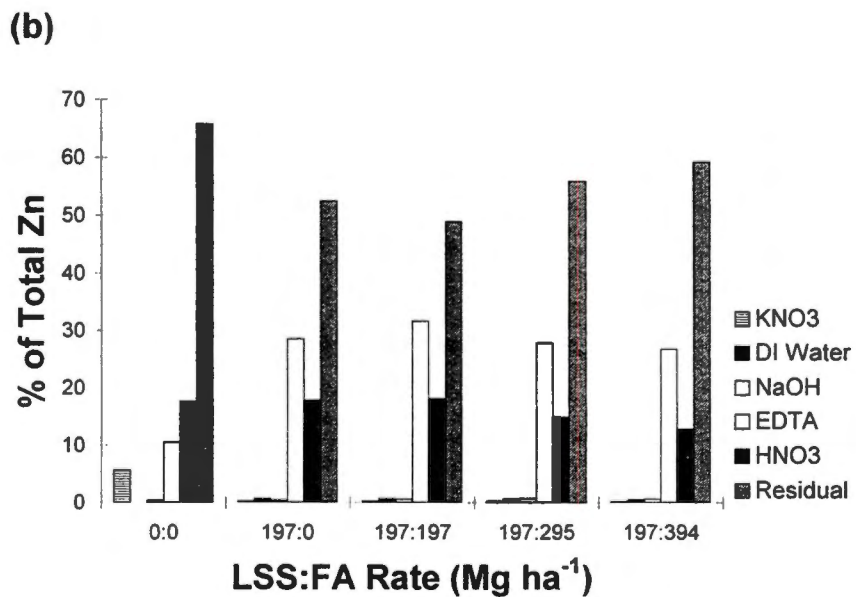
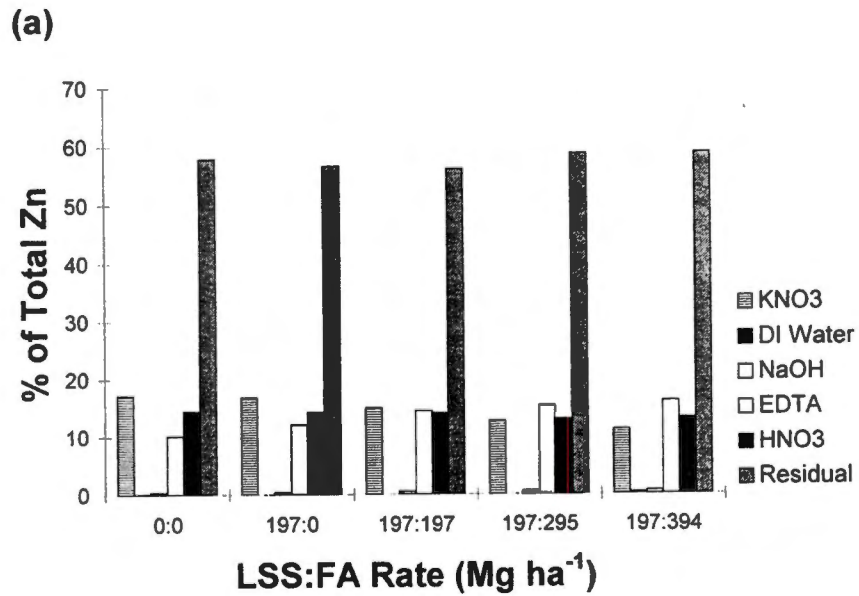


Figure 28. Influence of weathering and lime-stabilized sewage sludge (LSS) and fly ash (FA) amendments on the solid-phase speciation of zinc in the mine spoil after (a) 2 cycles and (b) 18 cycles of weathering.

exchangeable fraction. Copper speciation decreased in the soluble-exchangeable and residual fractions upon LSS application with an increase in the carbonate fraction. Primarily, weathering shifted Cu into the residual forms. Fly ash rate had no influence on Cu speciation for the weathered or unweathered mine spoil. In general, increasing FA rate increased Ni found in the residual fraction. Weathering shifted Ni into the carbonate and sulfide forms which decreased with FA application rate. However, Mn was primarily found in the soluble-exchangeable and residual fractions, irrespective of amendment. Weathering shifted Mn into the carbonate form in the amended mine spoil upon LSS application and into the residual form for the unamended mine spoil. Upon LSS application, Pb decreased in the soluble-exchangeable fraction. However, when weathered, amendment did not influence Pb solid-phase speciation. In general, FA rate increased the carbonate form of Zn in the unweathered material. However, weathering shifted Zn from the soluble-exchangeable and sulfide forms into the carbonate fraction, which was influenced by the LSS.

Chapter V

Summary and Conclusions

Summary

The objectives of this study were to (1) to examine the HNO₃-extractable elemental content of LSS, FA, and mine spoil and for comparison to total elemental content, (2) to evaluate the leachate chemistry of LSS and FA amended mine spoil during simulated laboratory weathering, and (3) to examine the solid-phase speciation of elements in LSS and FA amended, weathered and unweathered mine spoil. The clay mineralogy of the mine spoil material was identified as vermiculite, kaolinite, and mica. The bulk mineralogy of the unweathered mine spoil was primarily composed of quartz. The bulk mineralogy of the FA was dominated by quartz and mullite.

The total elemental content of the FA, LSS, and mine spoil was determined by employing both HNO₃ extraction and total digestion (aqua regia/HF). The percent recovery of elements by HNO₃ digestion was determined by dividing the HNO₃-extractable concentrations by the total elemental content determined using the aqua regia/HF method. Variations in HNO₃-extractability of metals in the mine spoil, LSS, and FA was a function of element speciation in the mineral phase. In general, HNO₃ digestion was more efficient at extracting Al, Ca, Mg, Cr, Cu, Pb, and Zn from

LSS; K and Na from the FA; and Fe, Mg, P, Co, and Mn from mine spoil. Further, HNO₃-extractability was not a function of total concentrations and the HNO₃ digestion method did not yield a constant percent extractability that could be uniformly applied, irrespective of material

Chemical analysis of leachate collected from simulated laboratory weathering showed that LSS neutralized solution pH for the duration of the study. The EC of all amended and unamended spoil material decreased over time, but the EC of the amended spoil material decreased less rapidly. Calcium and sulfate concentrations mirrored the EC of the solutions. High calcium concentrations were attributed to the LSS amendment. Sulfate concentrations decreased with time for all mine spoil material with greater sulfate concentrations in the amended mine spoil being attributed to LSS application, irrespective of FA treatment. Chloride and fluoride concentrations remained negligible throughout simulated weathering. Nitrate concentrations were did not vary with treatment, although fluctuation was observed during cycles 13 through 15. This fluctuation, irrespective of treatment, was attributed to increased microbial activity, and may have contributed to pH fluctuation during the same time period.

Leachate phosphorus concentrations increased with LSS application and FA treatment. A slight increase was observed as FA rate increased. Leachate potassium

concentrations in the unamended mine spoil increased with time, due to the dissolution of K-bearing minerals in the low pH system. In general, B concentrations increased with increasing FA rate, but concentrations were below 0.25 mg L^{-1} for all treatments. Leachate aluminum concentrations were higher for the unamended mine spoil material due to Al hydrolysis in low pH systems. Aluminum concentrations remained at or below detection limits for all amended mine spoil material.

Leachate copper concentrations were highest in the unamended mine spoil. This was attributed to increased mobility in low pH systems. Amended spoil material maintained leachate Cu concentrations at or below detection limits throughout the simulated weathering study, irrespective of FA treatment. Similarly leachate, Mn, Ni, and Zn concentrations were highest for the unamended mine spoil while the amended spoil material maintained a steady fluctuation at or below detectable levels throughout the study. This was attributed to reduced mobility due to LSS application. Leachate concentrations of Mn, Ni, and Zn were also unaffected by FA application rate.

In summary, the SSD procedure used in this study examined the solid-phase speciation of Ba, Co, Cr, Cu, Mn, Ni, Pb, Sr, and Zn. The speciation of Ba increased in the soluble-exchangeable and carbonate fractions with increasing FA rate. Strontium increased in the soluble-exchangeable

and sulfide fractions with increasing FA rate. Weathering tended to shift Ba to the carbonate fraction and Sr showed no significant shifts to any other phase. Chromium found in the residual fraction significantly increased upon LSS application. Weathering of the LSS-amended mine spoil increased Cr found in the adsorbed fraction, yet maintained predominance in the residual fraction. Chromium solid-phase speciation was not impacted by FA application or weathering. Although Co was primarily found in the residual fraction, Co showed a significant decrease in the soluble-exchangeable form upon LSS and FA application in the weathered material. In general, FA did not impact Co solid-phase speciation. However, LSS application decreased Co found in the soluble-exchangeable fraction. Copper speciation decreased in the soluble-exchangeable and residual fractions upon LSS application with an increase in the carbonate fraction. Primarily, weathering shifted Cu into the residual forms. Fly ash rate had no influence on Cu speciation for the weathered or unweathered mine spoil. In general, increasing FA rate increased Ni found in the residual fraction. Weathering shifted Ni into the carbonate and sulfide forms which decreased with FA application rate. However, Mn was primarily found in the soluble-exchangeable and residual fractions, irrespective of amendment. Weathering shifted Mn into the carbonate form in the amended mine spoil upon LSS application and into the residual form for the unamended

mine spoil. Upon LSS application, Pb decreased in the soluble-exchangeable fraction. However, when weathered, amendment did not influence Pb solid-phase speciation. In general, FA rate increased the carbonate form of Zn in the unweathered material. However, weathering shifted Zn from the soluble-exchangeable and sulfide forms into the carbonate fraction, which was influenced by the LSS.

Conclusions

The conclusions of this study are as follows:

- (1) HNO₃-extractability of the amendments was not a function of total concentrations and the HNO₃ digestion method did not yield a constant percent extractability that could be uniformly applied, irrespective of material.
- (2) LSS was determined to be beneficial to acid mine spoil reclamation by neutralizing solution pH and introducing essential nutrients.
- (3) LSS decreased heavy metal mobility and mineral dissolution by increasing the pH of the system.
- (4) In general, FA application enhanced phosphorus and boron concentrations.
- (5) In general, the impact of FA co-application with LSS was element specific with respect to the solid-phase speciation of elements.

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Appendices

Appendix A.
General Information

3 grams LSS= ~ 2.5 grams (minus water content of LSS)
 Acid base account (LSS)= 127.2
 $127.2 / 2.625 = X / 3$
 X= 145

Material	Neutralization Potential	Acid Potential	Acid-base Accounting
spoil	0	12.5	-12.5
sludge	131.89	4.69	127.2

to neutralize spoil:
 $0 = (145 * X) - 12.5$
 $12.5 = 145 * X$
 $X = .086$ (* 1000 tons of spoil)
 X = 86 tons of LSS/ 1000 tons spoil

Rounded to 88 tons LSS/ 1000 tons spoil for +- 0.002 error
 (water content = 12.5 %)

Acid-base accounting used to determine lime-stabilized sewage sludge application rate.

Notes:

1 hectare= 10,000 m²
 1 hectare (18 cm deep)= 10,000 m² * 0.18 m= 1800 m³
 1 hectare (18 cm deep) with bulk density of 1,300 kg/m³ weighs 2,340,000 kg
 1 hectare (18 cm deep)= 2,340,000 kg * 2.205 = 5,159,700 lbs
 1 acre= 2.471 hectares
 5,159,700 / 2.471= 2,088,101 lbs
 An accepted estimate is 1 acre= 2,000,000 lb= 1000 tons
 88 tons/1000 tons = 88 tons/ acre

Appendix B.
Amendment Characteristics

Amendment characteristics per HNO3 digestion				
elem	sam. name	conc(ppm)	final conc	average
Al	FA	256.284	25628.363	27917.436
Al	FA	302.065	30206.509	
As	FA	0.100	10.000	10.000
As	FA	0.100	10.000	
Ca	FA	67.838	6783.834	7414.751
Ca	FA	80.457	8045.667	
Cd	FA	0.015	1.539	1.264
Cd	FA	0.010	0.988	
Co	FA	0.226	22.606	23.830
Co	FA	0.251	25.054	
Cr	FA	0.149	14.882	18.153
Cr	FA	0.214	21.423	
Cu	FA	1.019	101.910	102.987
Cu	FA	1.041	104.063	
Fe	FA	89.743	8974.250	10416.487
Fe	FA	118.587	11858.723	
K	FA	57.467	5746.714	5879.171
K	FA	60.116	6011.627	
Mg	FA	22.492	2249.196	2355.717
Mg	FA	24.622	2462.238	
Mn	FA	0.759	75.940	82.001
Mn	FA	0.881	88.061	
Na	FA	5.899	589.944	622.569
Na	FA	6.552	655.193	
Ni	FA	0.372	37.223	43.355
Ni	FA	0.495	49.486	
P	FA	13.126	1312.639	1492.240
P	FA	16.718	1671.841	
Pb	FA	0.553	55.320	60.638
Pb	FA	0.660	65.955	
Se	FA	0.100	10.000	10.000
Se	FA	0.100	10.000	
Zn	FA	0.995	99.538	105.673
Zn	FA	1.118	111.807	
Al	spoil	86.965	86964.710	90025.655
Al	spoil	93.087	93086.600	
As	spoil	0.100	100.000	100.000
As	spoil	0.100	100.000	
Ca	spoil	18.988	18987.960	18600.915
Ca	spoil	18.214	18213.870	
Cd	spoil	0.004	4.000	4.000
Cd	spoil	0.004	4.000	
Co	spoil	0.115	115.380	112.690
Co	spoil	0.110	110.000	
Cr	spoil	0.163	163.220	164.150
Cr	spoil	0.165	165.080	
Cu	spoil	0.165	165.200	166.345
Cu	spoil	0.167	167.490	
Fe	spoil	498.138	498138.360	499069.180
Fe	spoil	500.000	500000.000	
K	spoil	16.395	16395.340	17209.295
K	spoil	18.023	18023.250	
Mg	spoil	22.353	22353.040	22183.795
Mg	spoil	22.015	22014.550	
Mn	spoil	2.994	2994.100	2777.880
Mn	spoil	2.562	2561.660	
Na	spoil	0.855	854.800	843.195
Na	spoil	0.832	831.590	
Ni	spoil	0.263	262.980	263.275
Ni	spoil	0.264	263.570	
P	spoil	13.642	13641.780	13643.490
P	spoil	13.645	13645.200	
Pb	spoil	0.271	271.310	278.875
Pb	spoil	0.286	286.440	
Se	spoil	0.101	100.700	116.925
Se	spoil	0.133	133.150	
Zn	spoil	0.939	939.290	929.995
Zn	spoil	0.921	920.700	
Al	SS	49.376	49376.330	50039.135
Al	SS	50.702	50701.940	
As	SS	0.100	100.000	100.000
As	SS	0.100	100.000	
Ca	SS	582.150	582149.840	581783.870
Ca	SS	581.418	581417.900	
Cd	SS	0.004	4.000	4.000
Cd	SS	0.004	4.000	
Co	SS	0.015	14.900	16.095

Co	SS	0.017	17.290	
Cr	SS	0.219	219.330	215.855
Cr	SS	0.212	212.380	
Cu	SS	0.623	623.280	637.235
Cu	SS	0.651	651.190	
Fe	SS	91.044	91044.370	90630.810
Fe	SS	90.217	90217.250	
K	SS	3.488	3488.370	3372.090
K	SS	3.256	3255.810	
Mg	SS	13.751	13750.680	14026.385
Mg	SS	14.302	14302.090	
Mn	SS	2.460	2460.370	2467.025
Mn	SS	2.474	2473.680	
Na	SS	2.216	2215.740	2195.840
Na	SS	2.176	2175.940	
Ni	SS	0.123	123.270	125.325
Ni	SS	0.127	127.380	
P	SS	45.159	45158.760	45486.725
P	SS	45.815	45814.690	
Pb	SS	0.201	201.490	195.285
Pb	SS	0.189	189.080	
Se	SS	0.100	100.000	100.000
Se	SS	0.100	100.000	
Zn	SS	1.770	1770.340	1777.830
Zn	SS	1.785	1785.320	

Amendment characteristics per totals analysis				
sample name	elem	conc (ppm)	conc w/dil	ave conc
FA	Al	204.109	102054.500	105876.538
FA	Al	215.275	107637.745	
FA	Al	215.875	107937.370	
FA	As	0.100	0.100	0.100
FA	As	0.100	0.100	
FA	As	0.100	0.100	
FA	Ba	1.803	901.520	5019.307
FA	Ba	26.428	13214.055	
FA	Ba	1.885	942.345	
FA	Ca	17.223	8611.565	18495.335
FA	Ca	76.108	38054.175	
FA	Ca	17.641	8820.265	
FA	Cd	0.009	4.430	7.687
FA	Cd	0.005	2.440	
FA	Cd	0.032	16.190	
FA	Co	0.105	52.300	58.493
FA	Co	0.115	57.455	
FA	Co	0.131	65.725	
FA	Cr	0.185	92.385	92.830
FA	Cr	0.182	91.050	
FA	Cr	0.190	95.055	
FA	Cu	0.369	184.250	193.828
FA	Cu	0.370	184.760	
FA	Cu	0.425	212.475	
FA	Fe	65.167	32583.255	32695.082
FA	Fe	63.561	31780.280	
FA	Fe	67.443	33721.710	
FA	K	12.067	6033.375	8023.105
FA	K	25.546	12772.785	
FA	K	10.526	5263.155	
FA	Mg	10.429	5214.530	10882.232
FA	Mg	44.122	22061.130	
FA	Mg	10.742	5371.035	
FA	Mn	0.271	135.600	234.932
FA	Mn	0.833	416.370	
FA	Mn	0.306	152.825	
FA	Mo	0.010	0.010	3.827
FA	Mo	0.010	5.190	
FA	Mo	0.013	6.280	
FA	Na	10.779	5389.635	31156.368
FA	Na	151.004	75502.135	
FA	Na	25.155	12577.335	
FA	Ni	0.162	80.785	81.778
FA	Ni	0.157	78.720	
FA	Ni	0.172	85.830	
FA	P	4.881	2440.415	2458.743
FA	P	4.756	2378.215	
FA	P	5.115	2557.600	
FA	Pb	0.425	212.650	226.068
FA	Pb	0.472	236.170	
FA	Pb	0.459	229.385	
FA	S	0.607	303.720	489.557
FA	S	1.756	877.995	
FA	S	0.574	286.955	
FA	Si	383.255	191627.300	270552.493
FA	Si	700.000	350000.000	
FA	Si	540.060	270030.180	
FA	Sr	1.477	738.710	811.558
FA	Sr	1.824	912.055	
FA	Sr	1.568	783.910	
FA	Ti	11.902	5950.985	6039.933
FA	Ti	11.986	5992.920	
FA	Ti	12.352	6175.895	
FA	Zn	0.334	167.145	176.490
FA	Zn	0.337	168.650	
FA	Zn	0.387	193.675	
FA	Zr	0.320	159.970	276.885
FA	Zr	0.401	200.685	
FA	Zr	0.940	470.000	
LSS	Al	30.842	15421.210	13358.772
LSS	Al	27.405	13702.460	
LSS	Al	21.905	10952.645	

LSS	As	0.100	0.100	0.100
LSS	As	0.100	0.100	
LSS	As	0.100	0.100	
LSS	Ba	12.023	6011.385	4033.457
LSS	Ba	11.937	5968.345	
LSS	Ba	0.241	120.640	
LSS	Ca	142.791	71395.635	63469.593
LSS	Ca	131.995	65997.715	
LSS	Ca	106.031	53015.460	
LSS	Cd	0.004	0.004	0.004
LSS	Cd	0.004	0.004	
LSS	Cd	0.004	0.004	
LSS	Co	0.013	6.640	6.030
LSS	Co	0.014	7.190	
LSS	Co	0.009	4.260	
LSS	Cr	0.051	25.640	22.862
LSS	Cr	0.044	21.760	
LSS	Cr	0.042	21.185	
LSS	Cu	0.173	86.505	75.210
LSS	Cu	0.121	60.605	
LSS	Cu	0.157	78.520	
LSS	Fe	24.449	12224.485	10943.055
LSS	Fe	20.495	10247.380	
LSS	Fe	20.715	10357.300	
LSS	K	10.112	5056.180	3910.467
LSS	K	9.628	4813.860	
LSS	K	3.723	1861.360	
LSS	Mg	20.246	10123.235	7251.903
LSS	Mg	19.934	9966.900	
LSS	Mg	3.331	1665.575	
LSS	Mn	0.998	499.095	415.100
LSS	Mn	0.974	487.230	
LSS	Mn	0.518	258.975	
LSS	Mo	0.010	0.010	0.010
LSS	Mo	0.010	0.010	
LSS	Mo	0.010	0.010	
LSS	Na	82.474	41237.045	27294.773
LSS	Na	74.572	37286.125	
LSS	Na	6.722	3361.150	
LSS	Ni	0.028	0.020	0.020
LSS	Ni	0.020	0.020	
LSS	Ni	0.020	0.020	
LSS	P	10.382	5190.830	4662.720
LSS	P	8.666	4333.030	
LSS	P	8.929	4464.300	
LSS	Pb	0.162	81.025	54.707
LSS	Pb	0.098	49.060	
LSS	Pb	0.068	34.035	
LSS	S	6.113	3056.450	2282.818
LSS	S	4.154	2076.865	
LSS	S	3.430	1715.140	
LSS	Si	207.147	103573.570	98109.310
LSS	Si	256.081	128040.585	
LSS	Si	125.428	62713.775	
LSS	Sr	0.241	120.560	86.025
LSS	Sr	0.188	93.975	
LSS	Sr	0.087	43.540	
LSS	Ti	0.994	496.870	415.025
LSS	Ti	0.850	425.000	
LSS	Ti	0.646	323.205	
LSS	Zn	0.388	193.855	180.043
LSS	Zn	0.355	177.455	
LSS	Zn	0.338	168.820	
LSS	Zr	0.143	71.315	43.797
LSS	Zr	0.074	37.205	
LSS	Zr	0.046	22.870	
SP	Al	153.994	76996.855	72739.142
SP	Al	146.885	73442.665	
SP	Al	135.556	67777.905	
SP	As	0.100	0.100	0.100
SP	As	0.100	0.100	
SP	As	0.100	0.100	
SP	Ba	28.833	14416.620	5075.258
SP	Ba	0.815	407.315	

SP	Ba	0.804	401.840	
SP	Ca	73.396	36698.125	14731.453
SP	Ca	7.791	3895.310	
SP	Ca	7.202	3600.925	
SP	Cd	0.004	0.004	0.004
SP	Cd	0.004	0.004	
SP	Cd	0.021	10.385	
SP	Co	0.039	19.715	17.267
SP	Co	0.023	11.675	
SP	Co	0.041	20.410	
SP	Cr	0.135	67.380	65.915
SP	Cr	0.134	66.805	
SP	Cr	0.127	63.560	
SP	Cu	0.035	17.715	22.587
SP	Cu	0.046	22.935	
SP	Cu	0.054	27.110	
SP	Fe	91.459	45729.670	44840.942
SP	Fe	90.951	45475.525	
SP	Fe	86.635	43317.630	
SP	K	25.931	12965.340	9798.885
SP	K	18.870	9435.170	
SP	K	13.992	6996.145	
SP	Mg	48.142	24070.770	11279.418
SP	Mg	9.963	4981.610	
SP	Mg	9.572	4785.875	
SP	Mn	1.221	610.745	333.163
SP	Mn	0.385	192.405	
SP	Mn	0.393	196.340	
SP	Mo	0.010	0.010	0.010
SP	Mo	0.011	0.010	
SP	Mo	0.010	0.010	
SP	Na	170.549	85274.425	40326.588
SP	Na	39.959	19979.585	
SP	Na	31.452	15725.755	
SP	Ni	0.020	0.020	0.020
SP	Ni	0.047	0.020	
SP	Ni	0.020	0.020	
SP	P	2.053	1026.725	989.875
SP	P	1.976	987.900	
SP	P	1.910	955.000	
SP	Pb	0.318	158.845	139.010
SP	Pb	0.269	134.720	
SP	Pb	0.247	123.465	
SP	S	16.532	8266.150	7657.773
SP	S	15.004	7502.215	
SP	S	14.410	7204.955	
SP	Si	700.000	350000.000	347186.045
SP	Si	700.000	350000.000	
SP	Si	683.116	341558.135	
SP	Sr	0.461	230.525	140.732
SP	Sr	0.184	92.045	
SP	Sr	0.199	99.625	
SP	Ti	4.403	2201.430	2040.528
SP	Ti	4.045	2022.400	
SP	Ti	3.796	1897.755	
SP	Zn	0.328	164.170	109.835
SP	Zn	0.162	80.755	
SP	Zn	0.169	84.580	
SP	Zr	0.234	117.115	220.205
SP	Zr	0.375	187.265	
SP	Zr	0.712	356.235	

Appendix C.
Simulated Laboratory Weathering

17	2	1	0.0500	0.0148	46.2105	0.0056	0.0100	2.3129	1.8959	0.0020	0.0200	0.0127	0.0080	0.7340	1.0740	2.9710	103.0250	7.49	0.24
17	2	2	0.0500	0.0131	33.7126	0.0040	0.0100	1.7322	1.4410	0.0020	0.0200	0.0379	0.0080	1.0430	1.2690	6.9980	75.9620	7.23	0.21
17	2	3	0.0500	0.0173	39.1796	0.0059	0.0100	2.5850	1.6063	0.0020	0.0200	0.0765	0.0080	0.7190	1.1200	1.2450	88.0300	7.42	0.22
17	3	1	0.0500	0.0748	0.0166	0.0100	2.6531	1.9434	0.0095	0.0200	0.0761	0.0080	0.9230	1.1950	4.3210	104.7990	7.08	0.24	
17	3	2	0.0500	0.0830	37.4514	0.0079	0.0100	2.7211	1.6877	0.0020	0.0200	0.0645	0.0080	0.7490	1.1180	9.9100	90.9350	7.11	0.22
17	3	3	0.0500	0.0854	41.7093	0.0071	0.0100	2.5850	1.7870	0.0020	0.0200	0.0808	0.0080	0.3810	1.1710	1.6590	84.9140	7.6	0.22
17	4	1	0.0500	0.1241	44.1145	0.0103	0.0100	2.4490	1.8851	0.0074	0.0200	0.0806	0.0080	0.3000	1.0360	1.0640	95.9630	7.63	0.24
17	4	2	0.0500	0.1002	42.8984	0.0091	0.0100	2.6531	1.8915	0.0106	0.0200	0.1023	0.0080	0.7770	1.1040	2.4450	102.1060	7.56	0.25
17	4	3	0.0500	0.0978	39.8780	0.0068	0.0100	2.3810	1.7422	0.0020	0.0200	0.0646	0.0080	0.9440	1.2450	3.1500	92.9390	7.4	0.22
17	5	1	0.0500	0.1199	35.5965	0.0071	0.0100	2.1088	1.6537	0.0020	0.0200	0.1424	0.0080	0.8820	1.2040	4.9660	81.5560	7.23	0.22
17	5	2	0.0500	0.1331	38.5995	0.0076	0.0100	2.4490	1.7044	0.0020	0.0200	0.1251	0.0080	0.8500	1.1970	6.4660	93.6720	7.27	0.22
17	5	3	0.0500	0.1208	40.4555	0.0082	0.0100	2.3129	1.6928	0.0020	0.0200	0.0548	0.0080	1.1240	1.3200	6.0770	94.1950	7.3	0.23
18	1	1	1.2704	0.0222	13.9253	0.0229	0.6504	6.8707	1.6794	0.9552	0.0200	0.0347	0.1425	0.5140	1.0030	1.4490	87.8500	3.61	0.24
18	1	2	1.0901	0.0181	11.7871	0.0244	0.6343	6.1224	1.5288	0.8114	0.0200	0.0100	0.1281	0.6490	0.9140	3.9600	79.9720	3.52	0.26
18	1	3	1.0161	0.0177	10.7718	0.0089	0.0772	5.7295	1.3524	0.7040	0.0200	0.0100	0.0864	0.1450	0.8130	2.6430	79.5870	3.52	0.24
18	2	1	0.0500	0.0164	42.2807	0.0074	0.0100	2.7211	1.7845	0.0025	0.0200	0.0937	0.0080	0.2740	0.8590	4.6430	88.1200	7.53	0.25
18	2	2	0.0500	0.0123	34.5115	0.0040	0.0100	1.7322	1.4578	0.0020	0.0200	0.0974	0.0080	0.2930	1.0840	2.5860	68.7330	7.45	0.21
18	2	3	0.0500	0.0177	33.8032	0.0040	0.0100	1.6656	1.4048	0.0020	0.0200	0.0786	0.0080	0.7510	1.2460	1.2600	72.6100	7.06	0.21
18	3	1	0.0500	0.0748	36.3066	0.0079	0.0100	2.7211	1.6839	0.0020	0.0200	0.1056	0.0080	0.5360	1.1480	5.5000	85.4160	7	0.23
18	3	2	0.0500	0.0668	28.2132	0.0040	0.0100	2.0653	1.3091	0.0020	0.0200	0.0853	0.0080	0.5640	1.0990	1.5420	72.0050	7.51	0.19
18	3	3	0.0500	0.0904	38.7081	0.0085	0.0100	2.4490	1.6678	0.0020	0.0200	0.0991	0.0080	0.6880	1.1770	1.3590	76.4580	7.62	0.22
18	4	1	0.0500	0.1208	37.0110	0.0106	0.0100	2.5170	1.6018	0.0020	0.0200	0.0904	0.0080	0.7610	1.2680	1.4420	81.1690	7.5	0.22
18	4	2	0.0500	0.1084	43.3059	0.0103	0.0100	2.5170	1.9017	0.0113	0.0200	0.1758	0.0080	0.3200	1.0160	1.5410	87.0570	7.8	0.24
18	4	3	0.0500	0.1084	32.0598	0.0094	0.0100	2.5170	1.4481	0.0062	0.0200	0.1391	0.0080	0.3620	1.0100	9.1080	80.9510	6.91	0.2
18	5	1	0.0500	0.1175	30.4010	0.0040	0.0100	1.3991	1.4687	0.0020	0.0200	0.1490	0.0080	0.6300	1.2190	6.7740	74.5910	7.18	0.2
18	5	2	0.0500	0.1249	32.9325	0.0203	0.0100	2.6531	1.4993	0.0143	0.0200	0.1721	0.0102	0.4570	1.1480	2.3640	71.7060	7.47	0.2
18	5	3	0.0500	0.1183	36.3845	0.0040	0.0100	1.8654	1.5951	0.0020	0.0200	0.1042	0.0080	0.4570	1.1920	3.5680	80.3090	7.47	0.23

Appendix D.
HNO₃ and Aqua Regia/HF Analyses

2 week weathering cycle: HNO3 digestion				TRT1
Element	Container	Concentration	Triplicate average	Replicate average
			mg kg-1	
Al	2	9490.21	9880.16	9828.10
Al	2	10081.03		
Al	2	10069.24		
Al	5	10954.57	10258.19	
Al	5	9527.25		
Al	5	10292.75		
Al	9	9420.63	9345.95	
Al	9	10394.99		
Al	9	8222.23		
Ca	2	858.66	994.05	1174.69
Ca	2	1277.21		
Ca	2	846.27		
Ca	5	1076.98	1184.96	
Ca	5	1090.49		
Ca	5	1387.42		
Ca	9	1723.57	1345.05	
Ca	9	1055.12		
Ca	9	1256.47		
Cd	2	0.41	0.62	0.70
Cd	2	0.99		
Cd	2	0.47		
Cd	5	0.40	0.77	
Cd	5	1.33		
Cd	5	0.59		
Cd	9	0.56	0.69	
Cd	9	1.10		
Cd	9	0.40		
Co	2	15.29	12.27	11.45
Co	2	11.02		
Co	2	10.49		
Co	5	10.26	11.21	
Co	5	12.26		
Co	5	11.10		
Co	9	11.01	10.88	
Co	9	11.30		
Co	9	10.34		
Cr	2	14.88	19.18	17.22
Cr	2	25.23		
Cr	2	17.43		
Cr	5	19.07	17.55	
Cr	5	15.76		
Cr	5	17.81		
Cr	9	15.19	14.92	
Cr	9	16.32		
Cr	9	13.23		
Cu	2	18.52	17.73	17.89
Cu	2	18.02		
Cu	2	16.65		
Cu	5	16.82	17.50	
Cu	5	18.96		
Cu	5	16.74		
Cu	9	17.99	18.44	
Cu	9	19.01		
Cu	9	18.33		
Fe	2	44995.29	47808.43	48807.95
Fe	2	50000.00		
Fe	2	48429.99		
Fe	5	50000.00	49949.10	
Fe	5	49847.30		
Fe	5	50000.00		
Fe	9	49741.95	48666.32	
Fe	9	46257.01		
Fe	9	50000.00		
K	2	2222.22	2119.24	2229.34
K	2	2031.43		
K	2	2104.07		
K	5	2531.07	2394.33	
K	5	2389.49		
K	5	2262.44		
K	9	2102.75	2174.43	
K	9	2544.80		

K	9	1875.75		
Mg	2	2155.32	2117.22	2126.71
Mg	2	2092.27		
Mg	2	2104.06		
Mg	5	2165.17	2176.08	
Mg	5	2222.61		
Mg	5	2140.46		
Mg	9	2090.17	2086.83	
Mg	9	2217.74		
Mg	9	1952.58		
Mn	2	396.43	270.37	233.86
Mn	2	196.58		
Mn	2	218.09		
Mn	5	251.50	249.65	
Mn	5	211.98		
Mn	5	285.48		
Mn	9	182.33	181.56	
Mn	9	188.61		
Mn	9	173.73		
Na	2	124.29	131.79	119.01
Na	2	149.18		
Na	2	121.88		
Na	5	111.02	107.85	
Na	5	103.79		
Na	5	108.75		
Na	9	112.74	117.39	
Na	9	139.96		
Na	9	99.47		
Ni	2	23.56	24.63	22.62
Ni	2	27.02		
Ni	2	23.30		
Ni	5	25.66	24.14	
Ni	5	22.61		
Ni	5	24.16		
Ni	9	19.59	19.11	
Ni	9	19.46		
Ni	9	18.27		
P	2	1197.37	1389.28	1409.90
P	2	1738.27		
P	2	1232.20		
P	5	1410.40	1533.72	
P	5	1328.08		
P	5	1862.68		
P	9	1352.92	1306.70	
P	9	1239.27		
P	9	1327.90		
Pb	2	30.68	27.63	28.79
Pb	2	24.06		
Pb	2	28.15		
Pb	5	29.89	31.38	
Pb	5	35.27		
Pb	5	28.99		
Pb	9	24.02	27.37	
Pb	9	30.73		
Pb	9	27.37		
Zn	2	87.74	87.22	89.40
Zn	2	87.46		
Zn	2	86.47		
Zn	5	91.68	90.90	
Zn	5	93.91		
Zn	5	87.10		
Zn	9	93.67	90.09	
Zn	9	89.32		
Zn	9	87.27		

Element	2 week weathering cycle: HNO3 digestion			TRT2
	Container	Concentration	Triplicate average	Replicate average
			mg kg-1	
Al	3	10332.02	10920.63	10402.24
Al	3	10385.78		
Al	3	12044.08		
Al	4	9745.49	10041.29	
Al	4	10335.58		
Al	4	10042.78		
Al	7	10925.64	10244.82	
Al	7	10123.02		
Al	7	9685.79		
Ca	3	3665.95	3916.59	3968.19
Ca	3	3897.86		
Ca	3	4185.95		
Ca	4	3483.90	3511.17	
Ca	4	3589.46		
Ca	4	3460.14		
Ca	7	5700.68	4476.82	
Ca	7	3372.52		
Ca	7	4357.26		
Cd	3	0.54	0.65	0.60
Cd	3	1.00		
Cd	3	0.40		
Cd	4	0.40	0.62	
Cd	4	1.06		
Cd	4	0.40		
Cd	7	0.65	0.55	
Cd	7	0.40		
Cd	7	0.59		
Co	3	9.80	10.61	11.25
Co	3	11.33		
Co	3	10.70		
Co	4	10.38	12.66	
Co	4	11.82		
Co	4	15.78		
Co	7	10.51	10.48	
Co	7	9.71		
Co	7	11.23		
Cr	3	17.15	17.57	17.72
Cr	3	18.93		
Cr	3	16.64		
Cr	4	19.18	18.53	
Cr	4	18.88		
Cr	4	17.53		
Cr	7	17.77	17.04	
Cr	7	17.41		
Cr	7	15.96		
Cu	3	21.39	21.85	21.48
Cu	3	22.17		
Cu	3	22.00		
Cu	4	20.96	21.01	
Cu	4	22.12		
Cu	4	19.93		
Cu	7	21.98	21.57	
Cu	7	19.57		
Cu	7	23.17		
Fe	3	49155.05	49718.35	49190.77
Fe	3	50000.00		
Fe	3	50000.00		
Fe	4	50000.00	50000.00	
Fe	4	50000.00		
Fe	4	50000.00		
Fe	7	47642.97	47853.97	
Fe	7	50000.00		
Fe	7	45918.94		
K	3	2154.88	2082.58	2129.22
K	3	2171.95		
K	3	1920.90		
K	4	2036.20	2080.66	
K	4	2115.38		
K	4	2090.40		
K	7	2278.34	2224.43	
K	7	2112.99		

K	7	2281.96		
Mg	3	2086.65	2116.73	2141.41
Mg	3	2141.32		
Mg	3	2122.22		
Mg	4	2137.16	2132.52	
Mg	4	2160.38		
Mg	4	2100.01		
Mg	7	2171.50	2174.99	
Mg	7	2127.45		
Mg	7	2226.02		
Mn	3	230.61	232.33	246.63
Mn	3	192.90		
Mn	3	273.49		
Mn	4	212.21	288.69	
Mn	4	263.40		
Mn	4	390.46		
Mn	7	199.67	218.87	
Mn	7	240.99		
Mn	7	215.94		
Na	3	103.83	103.70	100.20
Na	3	106.04		
Na	3	101.24		
Na	4	98.16	104.74	
Na	4	103.41		
Na	4	112.65		
Na	7	95.29	92.15	
Na	7	89.07		
Na	7	92.10		
Ni	3	22.74	23.88	24.13
Ni	3	24.53		
Ni	3	24.37		
Ni	4	25.06	25.43	
Ni	4	27.56		
Ni	4	23.67		
Ni	7	24.12	23.09	
Ni	7	24.43		
Ni	7	20.72		
P	3	1526.72	1580.08	1613.99
P	3	1569.92		
P	3	1643.59		
P	4	1541.03	1565.54	
P	4	1589.25		
P	4	1566.33		
P	7	2113.79	1696.34	
P	7	1468.41		
P	7	1506.83		
Pb	3	27.03	27.58	29.90
Pb	3	28.53		
Pb	3	27.19		
Pb	4	40.68	33.34	
Pb	4	29.14		
Pb	4	30.19		
Pb	7	27.38	28.77	
Pb	7	30.82		
Pb	7	28.13		
Zn	3	89.09	95.14	96.51
Zn	3	96.92		
Zn	3	99.41		
Zn	4	99.23	98.68	
Zn	4	103.93		
Zn	4	92.87		
Zn	7	93.96	95.70	
Zn	7	92.58		
Zn	7	100.55		

2 week weathering cycle: HNO3 digestion				TRT3
Element	Container	Concentration	Triplicate average mg kg ⁻¹	Replicate average
Al	6	11877.91	11514.46	11227.60
Al	6	11359.86		
Al	6	11305.63		
Al	14	11003.93	11016.54	
Al	14	10151.81		
Al	14	11893.90		
Al	15	11798.24	11151.79	
Al	15	11051.57		
Al	15	10605.57		
Ca	6	4260.77	4489.32	4006.49
Ca	6	5680.67		
Ca	6	3526.51		
Ca	14	3266.22	3588.42	
Ca	14	3746.94		
Ca	14	3752.10		
Ca	15	4267.68	3941.74	
Ca	15	4372.77		
Ca	15	3184.76		
Cd	6	0.40	0.55	0.64
Cd	6	0.61		
Cd	6	0.64		
Cd	14	0.40	0.78	
Cd	14	0.64		
Cd	14	1.30		
Cd	15	0.67	0.58	
Cd	15	0.51		
Cd	15	0.56		
Co	6	12.48	12.30	11.89
Co	6	11.80		
Co	6	12.63		
Co	14	11.48	12.22	
Co	14	12.56		
Co	14	12.64		
Co	15	12.16	11.13	
Co	15	10.60		
Co	15	10.64		
Cr	6	16.48	17.23	16.67
Cr	6	18.74		
Cr	6	16.48		
Cr	14	14.26	15.55	
Cr	14	14.88		
Cr	14	17.51		
Cr	15	18.46	17.22	
Cr	15	17.81		
Cr	15	15.40		
Cu	6	32.79	31.49	29.09
Cu	6	29.95		
Cu	6	31.73		
Cu	14	27.58	28.29	
Cu	14	28.59		
Cu	14	28.72		
Cu	15	28.13	27.50	
Cu	15	27.37		
Cu	15	26.99		
Fe	6	43290.30	45550.77	45557.11
Fe	6	48174.98		
Fe	6	45187.03		
Fe	14	47433.31	46364.53	
Fe	14	43675.09		
Fe	14	47985.19		
Fe	15	50000.00	44756.03	
Fe	15	42349.62		
Fe	15	41918.46		
K	6	2795.70	2628.31	2376.88
K	6	2413.02		
K	6	2676.22		
K	14	2289.56	2328.42	
K	14	2293.91		
K	14	2401.80		
K	15	2319.00	2173.92	
K	15	2092.76		

K	15	2109.99		
Mg	6	2330.41	2207.45	2148.88
Mg	6	2135.46		
Mg	6	2156.49		
Mg	14	2087.92	2132.74	
Mg	14	2118.61		
Mg	14	2191.69		
Mg	15	2187.33	2106.44	
Mg	15	2139.24		
Mg	15	1992.76		
Mn	6	213.20	214.66	211.40
Mn	6	242.69		
Mn	6	188.10		
Mn	14	177.24	199.97	
Mn	14	202.16		
Mn	14	220.50		
Mn	15	275.21	219.58	
Mn	15	196.01		
Mn	15	187.54		
Na	6	156.06	147.57	142.51
Na	6	141.71		
Na	6	144.95		
Na	14	136.61	134.91	
Na	14	135.25		
Na	14	132.86		
Na	15	146.10	145.08	
Na	15	141.36		
Na	15	147.71		
Ni	6	23.62	23.58	25.08
Ni	6	24.37		
Ni	6	22.74		
Ni	14	26.08	25.45	
Ni	14	26.33		
Ni	14	23.93		
Ni	15	26.82	26.22	
Ni	15	26.98		
Ni	15	24.88		
P	6	1478.83	1525.92	1539.98
P	6	1671.60		
P	6	1427.34		
P	14	1572.21	1547.04	
P	14	1431.29		
P	14	1637.63		
P	15	1717.89	1546.97	
P	15	1534.97		
P	15	1388.06		
Pb	6	34.41	33.33	31.08
Pb	6	32.43		
Pb	6	33.16		
Pb	14	32.84	31.42	
Pb	14	31.79		
Pb	14	29.64		
Pb	15	30.07	28.47	
Pb	15	30.33		
Pb	15	25.00		
Zn	6	100.10	99.35	96.63
Zn	6	98.43		
Zn	6	99.51		
Zn	14	95.46	97.94	
Zn	14	102.47		
Zn	14	95.91		
Zn	15	98.11	92.60	
Zn	15	94.44		
Zn	15	85.26		

Element	2 week weathering cycle: HNO3 digestion			TRT4
	Container	Concentration	Triplicate average mg kg-1	Replicate average
Al	8	11470.80	11493.65	12058.18
Al	8	11442.54		
Al	8	11567.62		
Al	10	12384.42	12455.14	
Al	10	12468.56		
Al	10	12502.43		
Al	13	12992.38	12225.75	
Al	13	11957.79		
Al	13	11727.06		
Ca	8	5186.51	5418.45	5132.89
Ca	8	7523.38		
Ca	8	3545.45		
Ca	10	5285.21	4739.47	
Ca	10	4508.61		
Ca	10	4424.59		
Ca	13	7306.65	5240.76	
Ca	13	3240.11		
Ca	13	5175.50		
Cd	8	0.62	0.61	0.65
Cd	8	0.69		
Cd	8	0.52		
Cd	10	0.47	0.68	
Cd	10	0.82		
Cd	10	0.74		
Cd	13	0.90	0.67	
Cd	13	0.61		
Cd	13	0.51		
Co	8	14.04	12.61	12.66
Co	8	12.19		
Co	8	11.61		
Co	10	12.10	13.19	
Co	10	13.22		
Co	10	14.26		
Co	13	12.63	12.17	
Co	13	11.54		
Co	13	12.34		
Cr	8	17.20	16.62	17.12
Cr	8	16.74		
Cr	8	15.91		
Cr	10	19.40	17.52	
Cr	10	16.02		
Cr	10	17.15		
Cr	13	18.33	17.23	
Cr	13	17.15		
Cr	13	16.22		
Cu	8	33.08	31.93	32.93
Cu	8	32.89		
Cu	8	29.81		
Cu	10	31.56	33.96	
Cu	10	35.16		
Cu	10	35.16		
Cu	13	34.82	32.89	
Cu	13	30.95		
Cu	13	32.89		
Fe	8	40703.74	41252.35	42528.53
Fe	8	39600.81		
Fe	8	43452.48		
Fe	10	45251.04	43921.78	
Fe	10	42853.46		
Fe	10	43660.83		
Fe	13	42576.87	42411.48	
Fe	13	44766.25		
Fe	13	39891.33		
K	8	2473.12	2397.21	2585.39
K	8	2485.07		
K	8	2233.45		
K	10	2488.69	2741.15	
K	10	2771.80		
K	10	2962.96		
K	13	2843.49	2617.81	
K	13	2345.68		

K	13	2664.28		
Mg	8	2196.32	2141.85	2182.35
Mg	8	2218.91		
Mg	8	2010.32		
Mg	10	2240.45	2263.84	
Mg	10	2313.27		
Mg	10	2237.80		
Mg	13	2296.04	2171.36	
Mg	13	1998.73		
Mg	13	2219.30		
Mn	8	234.50	205.90	216.44
Mn	8	193.18		
Mn	8	190.01		
Mn	10	203.23	249.03	
Mn	10	242.09		
Mn	10	301.76		
Mn	13	207.07	194.40	
Mn	13	185.32		
Mn	13	190.82		
Na	8	160.81	179.44	180.48
Na	8	192.55		
Na	8	184.97		
Na	10	174.39	174.83	
Na	10	173.84		
Na	10	176.26		
Na	13	192.81	187.16	
Na	13	205.99		
Na	13	162.69		
Ni	8	23.75	24.06	24.97
Ni	8	23.87		
Ni	8	24.56		
Ni	10	27.19	26.57	
Ni	10	25.76		
Ni	10	26.77		
Ni	13	22.61	24.27	
Ni	13	25.76		
Ni	13	24.44		
P	8	1538.75	1660.66	1606.83
P	8	1832.43		
P	8	1610.82		
P	10	1662.64	1577.27	
P	10	1531.81		
P	10	1537.36		
P	13	1752.94	1582.56	
P	13	1488.14		
P	13	1506.59		
Pb	8	38.97	33.87	33.34
Pb	8	34.36		
Pb	8	28.27		
Pb	10	32.30	33.71	
Pb	10	30.29		
Pb	10	38.55		
Pb	13	32.46	32.45	
Pb	13	30.54		
Pb	13	34.36		
Zn	8	111.45	100.94	100.96
Zn	8	104.50		
Zn	8	86.86		
Zn	10	101.62	103.83	
Zn	10	103.78		
Zn	10	106.11		
Zn	13	106.51	98.11	
Zn	13	88.49		
Zn	13	99.32		

2 week weathering cycle: HNO ₃ digestion				TRT5
Element	Container	Concentration	Triplicate average	Replicate average
			mg kg ⁻¹	
Al	1	11729.54		
Al	1	12643.68		
Al	1	13052.85		
Al	11	11551.58	12331.20	
Al	11	13506.65		
Al	11	11935.37		
Al	12	13398.30	12771.42	
Al	12	12443.63		
Al	12	12472.35		
Ca	1	4335.52	5402.25	5817.43
Ca	1	4573.44		
Ca	1	7297.80		
Ca	11	4284.99	5101.16	
Ca	11	5595.94		
Ca	11	5422.54		
Ca	12	6382.14	6948.89	
Ca	12	10200.36		
Ca	12	4264.16		
Cd	1	0.64	0.67	0.64
Cd	1	0.67		
Cd	1	0.70		
Cd	11	0.90	0.61	
Cd	11	0.40		
Cd	11	0.53		
Cd	12	0.40	0.65	
Cd	12	0.80		
Cd	12	0.74		
Co	1	13.07	12.57	13.31
Co	1	12.27		
Co	1	12.38		
Co	11	17.58	14.07	
Co	11	12.41		
Co	11	12.22		
Co	12	12.35	13.29	
Co	12	14.62		
Co	12	12.89		
Cr	1	16.02	18.11	17.83
Cr	1	19.57		
Cr	1	18.74		
Cr	11	15.65	17.76	
Cr	11	19.50		
Cr	11	18.11		
Cr	12	19.12	17.62	
Cr	12	17.77		
Cr	12	15.96		
Cu	1	33.91	34.71	35.48
Cu	1	32.36		
Cu	1	37.87		
Cu	11	35.65	35.18	
Cu	11	35.89		
Cu	11	34.01		
Cu	12	35.98	36.53	
Cu	12	38.26		
Cu	12	35.36		
Fe	1	42617.08	42339.10	43558.82
Fe	1	43657.64		
Fe	1	40742.58		
Fe	11	39386.51	43231.35	
Fe	11	44377.19		
Fe	11	45930.35		
Fe	12	46055.68	45106.00	
Fe	12	43726.17		
Fe	12	45536.17		
K	1	2473.12	2529.18	2568.85
K	1	2420.81		
K	1	2693.60		
K	11	2473.12	2452.30	
K	11	2734.46		
K	11	2149.32		
K	12	2836.16	2725.07	
K	12	2724.01		

K	12	2615.04		
Mg	1	2055.90	2133.22	2196.84
Mg	1	2119.75		
Mg	1	2224.01		
Mg	11	2074.98	2119.59	
Mg	11	2228.84		
Mg	11	2054.85		
Mg	12	2244.01	2337.70	
Mg	12	2669.49		
Mg	12	2099.60		
Mn	1	224.49	215.99	246.22
Mn	1	204.32		
Mn	1	219.16		
Mn	11	352.44	261.26	
Mn	11	217.19		
Mn	11	214.15		
Mn	12	203.24	261.40	
Mn	12	320.47		
Mn	12	260.50		
Na	1	209.77	205.15	194.46
Na	1	211.00		
Na	1	194.69		
Na	11	159.34	184.40	
Na	11	211.69		
Na	11	182.16		
Na	12	198.46	193.84	
Na	12	182.07		
Na	12	200.97		
Ni	1	25.32	25.52	27.14
Ni	1	25.86		
Ni	1	25.38		
Ni	11	27.78	26.50	
Ni	11	26.57		
Ni	11	25.17		
Ni	12	30.22	29.41	
Ni	12	30.11		
Ni	12	27.90		
P	1	1623.13	1703.44	1729.44
P	1	1742.85		
P	1	1744.32		
P	11	1629.34	1664.61	
P	11	1690.30		
P	11	1674.19		
P	12	1781.19	1820.26	
P	12	1947.18		
P	12	1732.42		
Pb	1	31.54	32.09	32.96
Pb	1	32.35		
Pb	1	32.37		
Pb	11	33.64	32.99	
Pb	11	34.07		
Pb	11	31.26		
Pb	12	34.98	33.81	
Pb	12	35.82		
Pb	12	30.63		
Zn	1	101.43	99.03	99.96
Zn	1	95.61		
Zn	1	100.05		
Zn	11	91.85	95.41	
Zn	11	98.04		
Zn	11	96.35		
Zn	12	102.40	105.44	
Zn	12	122.46		
Zn	12	91.46		

Element	Container	18 week weathering cycle: HNO3 digestion		mg kg-1
		Concentration	Triplicate average	
Al	2	10856.22		
Al	2	11180.46		
Al	2	10950.81		
Al	5	9789.30	9056.99	
Al	5	8321.69		
Al	5	9059.99		
Al	9	10313.74	10456.97	
Al	9	10847.07		
Al	9	10210.09		
Ca	2	366.09	416.36	378.95
Ca	2	406.03		
Ca	2	476.95		
Ca	5	399.22	350.02	
Ca	5	306.26		
Ca	5	344.58		
Ca	9	363.50	370.46	
Ca	9	352.25		
Ca	9	395.65		
Cd	2	0.40	0.40	0.40
Cd	2	0.40		
Cd	2	0.40		
Cd	5	0.40	0.41	
Cd	5	0.44		
Cd	5	0.40		
Cd	9	0.40	0.40	
Cd	9	0.40		
Cd	9	0.40		
Co	2	8.91	9.27	9.01
Co	2	9.03		
Co	2	9.86		
Co	5	8.49	8.62	
Co	5	8.42		
Co	5	8.94		
Co	9	8.91	9.15	
Co	9	9.09		
Co	9	9.45		
Cr	2	20.13	20.20	18.17
Cr	2	20.36		
Cr	2	20.13		
Cr	5	18.27	14.96	
Cr	5	12.98		
Cr	5	13.65		
Cr	9	19.01	19.35	
Cr	9	21.05		
Cr	9	17.99		
Cu	2	14.74	15.24	15.16
Cu	2	15.06		
Cu	2	15.93		
Cu	5	14.14	15.53	
Cu	5	16.25		
Cu	5	16.20		
Cu	9	14.69	14.70	
Cu	9	14.37		
Cu	9	15.06		
Fe	2	50000.00	50000.00	48220.48
Fe	2	50000.00		
Fe	2	50000.00		
Fe	5	50000.00	44861.45	
Fe	5	42440.44		
Fe	5	41543.92		
Fe	9	50000.00	50000.00	
Fe	9	50000.00		
Fe	9	50000.00		
K	2	2465.12	2480.62	2341.87
K	2	2511.63		
K	2	2465.12		
K	5	2139.53	2250.41	
K	5	2198.33		
K	5	2413.38		
K	9	2255.81	2294.57	
K	9	2418.60		

K	9	2209.30		
Mg	2	2129.12	2116.80	2016.62
Mg	2	2116.92		
Mg	2	2104.37		
Mg	5	1932.03	1906.97	
Mg	5	1877.11		
Mg	5	1911.77		
Mg	9	2020.20	2026.08	
Mg	9	2057.32		
Mg	9	2000.73		
Mn	2	139.93	148.35	155.89
Mn	2	149.82		
Mn	2	155.31		
Mn	5	138.83	134.87	
Mn	5	132.05		
Mn	5	133.72		
Mn	9	162.12	184.47	
Mn	9	146.50		
Mn	9	244.78		
Na	2	93.07	84.10	84.59
Na	2	67.18		
Na	2	92.06		
Na	5	68.60	90.39	
Na	5	96.32		
Na	5	106.24		
Na	9	84.61	79.27	
Na	9	65.87		
Na	9	87.34		
Ni	2	21.95	23.01	20.91
Ni	2	22.54		
Ni	2	24.54		
Ni	5	19.43	17.27	
Ni	5	15.68		
Ni	5	16.69		
Ni	9	22.31	22.44	
Ni	9	21.60		
Ni	9	23.42		
P	2	1342.67	1455.67	1351.86
P	2	1439.26		
P	2	1585.08		
P	5	1388.57	1222.44	
P	5	1130.03		
P	5	1148.71		
P	9	1351.87	1377.47	
P	9	1302.49		
P	9	1478.05		
Pb	2	34.31	33.87	29.01
Pb	2	31.86		
Pb	2	35.44		
Pb	5	25.96	24.13	
Pb	5	23.00		
Pb	5	23.44		
Pb	9	27.95	29.04	
Pb	9	31.78		
Pb	9	27.41		
Zn	2	82.90	86.96	81.34
Zn	2	86.59		
Zn	2	91.38		
Zn	5	77.76	73.49	
Zn	5	70.98		
Zn	5	71.72		
Zn	9	81.22	83.57	
Zn	9	83.77		
Zn	9	85.73		

18 week weathering cycle: HNO3 digestion				TRT2
Element	Container	Concentration	Triplicate average	Replicate average
			mg kg-1	
Al	3	10845.16	11742.68	10726.25
Al	3	11969.19		
Al	3	12413.70		
Al	4	9781.48	10286.71	
Al	4	9955.39		
Al	4	11123.25		
Al	7	10553.61	10149.37	
Al	7	10131.57		
Al	7	9762.91		
Ca	3	4375.97	5033.32	4827.00
Ca	3	5371.02		
Ca	3	5352.98		
Ca	4	4361.27	4685.96	
Ca	4	4695.74		
Ca	4	5000.87		
Ca	7	4757.74	4761.72	
Ca	7	5207.17		
Ca	7	4320.25		
Cd	3	0.44	0.41	0.42
Cd	3	0.40		
Cd	3	0.40		
Cd	4	0.56	0.45	
Cd	4	0.40		
Cd	4	0.40		
Cd	7	0.40	0.40	
Cd	7	0.40		
Cd	7	0.40		
Co	3	11.97	11.60	11.29
Co	3	10.70		
Co	3	12.13		
Co	4	11.38	11.22	
Co	4	11.12		
Co	4	11.18		
Co	7	10.58	11.04	
Co	7	11.24		
Co	7	11.30		
Cr	3	17.41	20.75	19.75
Cr	3	22.03		
Cr	3	22.81		
Cr	4	16.17	19.66	
Cr	4	20.31		
Cr	4	22.49		
Cr	7	21.05	18.85	
Cr	7	19.75		
Cr	7	15.76		
Cu	3	27.67	26.67	26.17
Cu	3	25.44		
Cu	3	26.91		
Cu	4	26.84	25.71	
Cu	4	24.53		
Cu	4	25.76		
Cu	7	25.26	26.13	
Cu	7	25.76		
Cu	7	27.38		
Fe	3	46349.28	48783.09	48291.40
Fe	3	50000.00		
Fe	3	50000.00		
Fe	4	44076.28	48025.43	
Fe	4	50000.00		
Fe	4	50000.00		
Fe	7	50000.00	48065.69	
Fe	7	50000.00		
Fe	7	44197.06		
K	3	2580.65	2530.76	2329.19
K	3	2453.49		
K	3	2558.14		
K	4	2329.75	2218.44	
K	4	2011.63		
K	4	2313.95		
K	7	2232.56	2238.36	
K	7	2093.02		

K	7	2389.49		
Mg	3	2246.37	2303.93	2194.05
Mg	3	2289.35		
Mg	3	2376.07		
Mg	4	2128.11	2147.95	
Mg	4	2095.45		
Mg	4	2222.29		
Mg	7	2140.49	2130.26	
Mg	7	2132.76		
Mg	7	2117.54		
Mn	3	307.83	315.49	284.68
Mn	3	302.76		
Mn	3	335.88		
Mn	4	267.77	270.44	
Mn	4	270.93		
Mn	4	272.62		
Mn	7	264.03	268.11	
Mn	7	290.37		
Mn	7	249.93		
Na	3	121.16	110.34	95.59
Na	3	110.55		
Na	3	99.30		
Na	4	77.70	85.29	
Na	4	88.63		
Na	4	89.55		
Na	7	85.99	91.14	
Na	7	89.18		
Na	7	96.26		
Ni	3	25.07	29.10	26.93
Ni	3	28.35		
Ni	3	33.87		
Ni	4	20.86	25.67	
Ni	4	27.47		
Ni	4	28.88		
Ni	7	26.94	26.03	
Ni	7	28.59		
Ni	7	22.55		
P	3	1731.97	1929.98	1880.62
P	3	2059.06		
P	3	1998.90		
P	4	1642.24	1865.75	
P	4	1876.67		
P	4	2078.33		
P	7	1877.49	1846.15	
P	7	2044.06		
P	7	1616.90		
Pb	3	32.02	34.84	32.66
Pb	3	33.92		
Pb	3	38.58		
Pb	4	27.70	31.27	
Pb	4	32.31		
Pb	4	33.80		
Pb	7	31.41	31.89	
Pb	7	37.37		
Pb	7	26.89		
Zn	3	113.20	120.70	116.04
Zn	3	119.09		
Zn	3	129.80		
Zn	4	106.94	113.82	
Zn	4	114.38		
Zn	4	120.14		
Zn	7	114.08	113.60	
Zn	7	119.65		
Zn	7	107.06		

18 week weathering cycle: HNO3 digestion				TRT3	
Element	Container	Concentration	Triplicate average	Replicate average	
			mg kg-1		
Al	6	12241.30	11996.67	12174.95	
Al	6	12140.99			
Al	6	11607.72			
Al	14	12093.03	11389.06		
Al	14	11445.78			
Al	14	10628.37			
Al	15	12747.14	13139.14		
Al	15	13284.84			
Al	15	13385.43			
Ca	6	5097.09	5063.94	5085.49	
Ca	6	5065.46			
Ca	6	5029.26			
Ca	14	4838.58	4803.45		
Ca	14	5385.94			
Ca	14	4185.84			
Ca	15	5274.60	5389.09		
Ca	15	5559.36			
Ca	15	5333.32			
Cd	6	0.40	0.40	0.41	
Cd	6	0.40			
Cd	6	0.40			
Cd	14	0.40	0.43		
Cd	14	0.40			
Cd	14	0.49			
Cd	15	0.40	0.40		
Cd	15	0.40			
Cd	15	0.40			
Co	6	12.14	12.06	12.14	
Co	6	12.08			
Co	6	11.96			
Co	14	12.61	12.04		
Co	14	11.78			
Co	14	11.74			
Co	15	12.91	12.32		
Co	15	11.90			
Co	15	12.14			
Cr	6	21.75	21.52	21.27	
Cr	6	21.24			
Cr	6	21.56			
Cr	14	21.33	19.47		
Cr	14	21.33			
Cr	14	15.76			
Cr	15	22.44	22.81		
Cr	15	22.77			
Cr	15	23.23			
Cu	6	31.44	31.06	31.35	
Cu	6	31.12			
Cu	6	30.62			
Cu	14	30.98	31.41		
Cu	14	30.94			
Cu	14	32.31			
Cu	15	31.16	31.58		
Cu	15	31.99			
Cu	15	31.58			
Fe	6	50000.00	50000.00	48848.74	
Fe	6	50000.00			
Fe	6	50000.00			
Fe	14	49402.50	46546.21		
Fe	14	49629.76			
Fe	14	40606.35			
Fe	15	50000.00	50000.00		
Fe	15	50000.00			
Fe	15	50000.00			
K	6	2500.00	2437.98	2506.27	
K	6	2441.86			
K	6	2372.09			
K	14	2523.26	2499.43		
K	14	2430.23			
K	14	2544.80			
K	15	2616.28	2581.40		
K	15	2616.28			

K	15	2511.63		
Mg	6	2232.30	2183.83	2187.48
Mg	6	2179.16		
Mg	6	2140.04		
Mg	14	2183.62	2131.27	
Mg	14	2136.76		
Mg	14	2073.42		
Mg	15	2231.76	2247.35	
Mg	15	2262.06		
Mg	15	2248.23		
Mn	6	260.42	263.69	286.18
Mn	6	278.65		
Mn	6	252.00		
Mn	14	429.89	299.41	
Mn	14	251.34		
Mn	14	217.01		
Mn	15	360.44	295.45	
Mn	15	264.17		
Mn	15	261.73		
Na	6	126.56	118.16	129.63
Na	6	105.57		
Na	6	122.34		
Na	14	125.36	132.31	
Na	14	131.92		
Na	14	139.65		
Na	15	131.07	138.43	
Na	15	141.25		
Na	15	142.96		
Ni	6	29.53	29.23	28.96
Ni	6	30.53		
Ni	6	27.65		
Ni	14	33.52	28.60	
Ni	14	29.41		
Ni	14	22.86		
Ni	15	28.71	29.06	
Ni	15	28.35		
Ni	15	30.11		
P	6	1947.36	1958.04	1943.22
P	6	1979.24		
P	6	1947.51		
P	14	1904.46	1851.21	
P	14	2059.24		
P	14	1589.92		
P	15	2055.95	2020.43	
P	15	2032.65		
P	15	1972.68		
Pb	6	35.71	35.11	34.99
Pb	6	33.20		
Pb	6	36.43		
Pb	14	36.39	33.66	
Pb	14	34.35		
Pb	14	30.24		
Pb	15	36.11	36.21	
Pb	15	36.46		
Pb	15	36.08		
Zn	6	116.12	115.63	115.46
Zn	6	115.58		
Zn	6	115.19		
Zn	14	115.92	111.99	
Zn	14	116.86		
Zn	14	103.19		
Zn	15	120.92	118.75	
Zn	15	118.12		
Zn	15	117.20		

18 week weathering cycle: HNO3 digestion				TRT4
Element	Container	Concentration	Triplicate average	Replicate average
			mg kg-1	
Al	8	12992.55	13419.00	12841.34
Al	8	12907.38		
Al	8	14357.07		
Al	10	12012.41	11992.08	
Al	10	13243.39		
Al	10	10720.43		
Al	13	13588.71	13112.93	
Al	13	13999.29		
Al	13	11750.79		
Ca	8	5452.94	5363.99	5171.51
Ca	8	5253.05		
Ca	8	5385.99		
Ca	10	5435.83	5428.74	
Ca	10	5849.53		
Ca	10	5000.85		
Ca	13	4821.66	4721.78	
Ca	13	4657.24		
Ca	13	4686.46		
Cd	8	0.40	0.45	0.64
Cd	8	0.40		
Cd	8	0.55		
Cd	10	0.40	0.55	
Cd	10	0.72		
Cd	10	0.54		
Cd	13	0.40	0.91	
Cd	13	0.40		
Cd	13	1.92		
Co	8	13.39	13.07	12.75
Co	8	12.38		
Co	8	13.45		
Co	10	12.32	12.36	
Co	10	13.45		
Co	10	11.30		
Co	13	12.73	12.82	
Co	13	12.08		
Co	13	13.67		
Cr	8	21.24	22.27	20.70
Cr	8	22.35		
Cr	8	23.23		
Cr	10	19.75	19.16	
Cr	10	22.58		
Cr	10	15.14		
Cr	13	22.77	20.67	
Cr	13	21.89		
Cr	13	17.35		
Cu	8	32.90	34.03	34.48
Cu	8	33.36		
Cu	8	35.83		
Cu	10	34.60	35.35	
Cu	10	36.15		
Cu	10	35.31		
Cu	13	33.50	34.07	
Cu	13	32.86		
Cu	13	35.84		
Fe	8	49933.00	49700.91	48903.74
Fe	8	49169.74		
Fe	8	50000.00		
Fe	10	44938.03	44533.04	
Fe	10	49855.35		
Fe	10	39005.75		
Fe	13	50000.00	46477.27	
Fe	13	48973.43		
Fe	13	40458.36		
K	8	2627.91	2693.80	2648.55
K	8	2604.65		
K	8	2848.84		
K	10	2372.09	2494.70	
K	10	2662.79		
K	10	2449.22		
K	13	2767.44	2757.16	
K	13	2744.19		

K	13	2759.86		
Mg	8	2138.76	2211.68	2198.13
Mg	8	2194.63		
Mg	8	2301.64		
Mg	10	2179.98	2189.76	
Mg	10	2324.75		
Mg	10	2064.56		
Mg	13	2249.23	2192.95	
Mg	13	2198.27		
Mg	13	2131.37		
Mn	8	333.44	289.53	265.52
Mn	8	256.89		
Mn	8	278.26		
Mn	10	255.98	251.79	
Mn	10	284.06		
Mn	10	215.33		
Mn	13	274.54	255.25	
Mn	13	265.08		
Mn	13	226.12		
Na	8	124.85	144.95	141.66
Na	8	151.62		
Na	8	158.37		
Na	10	158.69	143.91	
Na	10	143.97		
Na	10	129.07		
Na	13	125.46	136.11	
Na	13	152.33		
Na	13	130.54		
Ni	8	28.82	29.25	28.49
Ni	8	27.41		
Ni	8	31.52		
Ni	10	28.29	27.67	
Ni	10	31.23		
Ni	10	23.49		
Ni	13	32.17	28.53	
Ni	13	29.94		
Ni	13	23.49		
P	8	2131.30	2037.77	1920.53
P	8	2009.75		
P	8	1972.26		
P	10	1868.27	1879.64	
P	10	2057.94		
P	10	1712.72		
P	13	1950.17	1844.19	
P	13	1864.44		
P	13	1717.96		
Pb	8	30.53	34.91	34.89
Pb	8	37.37		
Pb	8	36.82		
Pb	10	34.78	34.45	
Pb	10	38.60		
Pb	10	29.96		
Pb	13	34.61	35.32	
Pb	13	36.89		
Pb	13	34.46		
Zn	8	112.11	114.95	113.43
Zn	8	112.51		
Zn	8	120.22		
Zn	10	112.13	113.45	
Zn	10	122.39		
Zn	10	105.81		
Zn	13	115.68	111.90	
Zn	13	110.72		
Zn	13	109.29		

18 week weathering cycle: HNO3 digestion				TRT5
Element	Container	Concentration	Triplicate average	Replicate average
			mg kg-1	
Al	1	15213.88	14489.95	13734.79
Al	1	14744.57		
Al	1	13451.40		
Al	11	14524.25	14430.59	
Al	11	14189.91		
Al	11	14577.61		
Al	12	12715.08	12303.84	
Al	12	12305.82		
Al	12	11890.61		
Ca	1	4796.65	4921.32	5024.45
Ca	1	4963.55		
Ca	1	5003.76		
Ca	11	4689.58	5026.08	
Ca	11	5058.97		
Ca	11	5329.70		
Ca	12	5149.06	5125.93	
Ca	12	5193.02		
Ca	12	5035.72		
Cd	1	0.40	0.40	0.45
Cd	1	0.40		
Cd	1	0.40		
Cd	11	0.87	0.56	
Cd	11	0.40		
Cd	11	0.40		
Cd	12	0.40	0.40	
Cd	12	0.40		
Cd	12	0.40		
Co	1	12.14	12.91	13.05
Co	1	13.93		
Co	1	12.67		
Co	11	12.99	13.54	
Co	11	13.45		
Co	11	14.17		
Co	12	12.61	12.71	
Co	12	12.67		
Co	12	12.85		
Cr	1	21.93	22.51	21.45
Cr	1	24.11		
Cr	1	21.47		
Cr	11	18.69	21.60	
Cr	11	22.58		
Cr	11	23.51		
Cr	12	20.73	20.26	
Cr	12	21.19		
Cr	12	18.87		
Cu	1	34.92	35.48	36.40
Cu	1	36.24		
Cu	1	35.28		
Cu	11	39.85	38.38	
Cu	11	37.34		
Cu	11	37.94		
Cu	12	35.92	35.34	
Cu	12	35.74		
Cu	12	34.37		
Fe	1	45716.66	47441.76	46547.73
Fe	1	50000.00		
Fe	1	46608.63		
Fe	11	40995.15	45239.63	
Fe	11	46887.84		
Fe	11	47835.89		
Fe	12	47744.21	46961.79	
Fe	12	47614.65		
Fe	12	45526.50		
K	1	2941.86	2899.22	2771.40
K	1	3011.63		
K	1	2744.19		
K	11	3070.49	2988.61	
K	11	2895.35		
K	11	3000.00		
K	12	2500.00	2426.36	
K	12	2418.60		

K	12	2360.47		
Mg	1	2238.03	2250.47	2233.05
Mg	1	2328.12		
Mg	1	2185.26		
Mg	11	2288.93	2322.64	
Mg	11	2315.92		
Mg	11	2363.06		
Mg	12	2175.98	2126.05	
Mg	12	2133.85		
Mg	12	2068.33		
Mn	1	215.77	240.20	258.09
Mn	1	255.23		
Mn	1	249.60		
Mn	11	232.29	256.72	
Mn	11	269.54		
Mn	11	268.32		
Mn	12	264.39	277.34	
Mn	12	260.85		
Mn	12	306.79		
Na	1	148.30	163.73	161.75
Na	1	168.55		
Na	1	174.33		
Na	11	154.79	171.94	
Na	11	175.72		
Na	11	185.30		
Na	12	132.83	149.58	
Na	12	161.86		
Na	12	154.03		
Ni	1	30.35	31.52	32.33
Ni	1	32.81		
Ni	1	31.41		
Ni	11	43.40	35.79	
Ni	11	32.64		
Ni	11	31.35		
Ni	12	29.94	29.68	
Ni	12	30.88		
Ni	12	28.24		
P	1	1875.10	1887.96	1883.15
P	1	1950.14		
P	1	1838.65		
P	11	1670.62	1830.35	
P	11	1889.78		
P	11	1930.64		
P	12	1980.53	1931.13	
P	12	1958.79		
P	12	1854.06		
Pb	1	37.15	36.57	36.09
Pb	1	37.51		
Pb	1	35.05		
Pb	11	35.51	37.01	
Pb	11	36.14		
Pb	11	39.38		
Pb	12	35.00	34.69	
Pb	12	35.26		
Pb	12	33.80		
Zn	1	109.24	115.30	114.69
Zn	1	118.15		
Zn	1	118.50		
Zn	11	112.58	115.91	
Zn	11	118.73		
Zn	11	118.41		
Zn	12	115.68	112.88	
Zn	12	113.98		
Zn	12	108.97		

Element	Container	2 week weathering cycle: totals (aqua regia/HF)		TRT1
		Concentration	Triplicate average	Replicate average
			mg kg-1	
2	Al	95533.25	81309.29	80952.82
2	Al	67742.11		
2	Al	80652.51		
5	Al	77837.20	85327.32	
5	Al	85106.66		
5	Al	93038.09		
9	Al	75008.16	76221.85	
9	Al	77883.41		
9	Al	75773.97		
2	Ba	430.81	415.02	449.87
2	Ba	471.26		
2	Ba	343.00		
5	Ba	436.74	480.74	
5	Ba	428.77		
5	Ba	576.72		
9	Ba		453.84	
9	Ba	453.84		
9	Ba			
2	Ca	553.54	577.96	3005.86
2	Ca	677.90		
2	Ca	502.44		
5	Ca	621.97	1584.84	
5	Ca	726.74		
5	Ca	3405.82		
9	Ca	10226.22	6854.79	
9	Ca	733.16		
9	Ca	9604.98		
2	Cd	2.00	2.00	2.00
2	Cd	2.00		
2	Cd	2.00		
5	Cd	2.00	2.00	
5	Cd	2.00		
5	Cd	2.00		
9	Cd	2.00	2.00	
9	Cd	2.00		
9	Cd	2.00		
2	Co	8.40	9.24	11.88
2	Co	10.26		
2	Co	9.06		
5	Co	9.69	12.64	
5	Co	8.63		
5	Co	19.62		
9	Co	12.11	13.77	
9	Co	17.16		
9	Co	12.05		
2	Cr	63.56	68.53	70.14
2	Cr	86.31		
2	Cr	55.74		
5	Cr	68.34	75.52	
5	Cr	78.80		
5	Cr	79.43		
9	Cr	64.33	66.36	
9	Cr	69.48		
9	Cr	65.28		
2	Cu	11.04	7.95	18.92
2	Cu	2.00		
2	Cu	10.81		
5	Cu	28.92	27.68	
5	Cu	6.03		
5	Cu	48.08		
9	Cu	19.59	21.14	
9	Cu	26.05		
9	Cu	17.77		
2	Fe	38515.39	41841.22	42284.93
2	Fe	52555.81		
2	Fe	34452.46		
5	Fe	41727.02	44463.99	
5	Fe	47631.36		
5	Fe	44033.59		
9	Fe	39525.73	40549.59	
9	Fe	42376.42		

9	Fe	39746.62		
2	K	4685.49	4440.97	5825.21
2	K	5107.25		
2	K	3530.17		
5	K	6611.04	5332.49	
5	K	5311.54		
5	K	4074.89		
9	K	7124.52	7702.18	
9	K	6803.60		
9	K	9178.43		
2	Mg	5194.51	5335.91	6644.71
2	Mg	6164.24		
2	Mg	4648.99		
5	Mg	5408.63	5611.08	
5	Mg	5483.81		
5	Mg	5940.79		
9	Mg	10858.12	8987.14	
9	Mg	5535.31		
9	Mg	10568.00		
2	Mn	125.41	129.97	169.31
2	Mn	152.15		
2	Mn	112.34		
5	Mn	138.42	142.79	
5	Mn	139.65		
5	Mn	150.29		
9	Mn	284.78	235.19	
9	Mn	148.26		
9	Mn	272.54		
2	Mo	5.00	5.06	5.02
2	Mo	5.00		
2	Mo	5.19		
5	Mo	5.00	5.00	
5	Mo	5.00		
5	Mo	5.00		
9	Mo	5.00	5.00	
9	Mo	5.00		
9	Mo	5.00		
2	Na	2302.91	5762.96	10625.08
2	Na	6285.54		
2	Na	8700.45		
5	Na	8944.48	7401.85	
5	Na	6973.89		
5	Na	6287.19		
9	Na	25685.39	18710.42	
9	Na	6339.47		
9	Na	24106.41		
2	Ni	10.00	15.22	15.35
2	Ni	25.66		
2	Ni	10.00		
5	Ni	10.00	18.86	
5	Ni	10.75		
5	Ni	35.84		
9	Ni	10.00	11.98	
9	Ni	14.46		
9	Ni	11.48		
2	P	708.32	773.51	839.68
2	P	996.94		
2	P	615.28		
5	P	731.76	1010.51	
5	P	883.10		
5	P	1416.67		
9	P	707.25	735.02	
9	P	746.27		
9	P	751.54		
2	Pb	120.06	124.98	137.05
2	Pb	132.37		
2	Pb	122.50		
5	Pb	132.53	137.62	
5	Pb	121.14		
5	Pb	159.18		
9	Pb	155.82	148.57	
9	Pb	123.23		
9	Pb	166.67		
2	S	3656.28	4781.77	4613.27

2	S	6968.24		
2	S	3720.79		
5	S	4024.64	5082.51	
5	S	6458.66		
5	S	4764.23		
9	S	4013.14	3975.54	
9	S	3924.14		
9	S	3989.35		
2	Si	155560.53	184795.43	201781.36
2	Si	191440.11		
2	Si	207385.67		
5	Si	264862.70	205729.07	
5	Si	182932.88		
5	Si	169391.65		
9	Si	204806.56	214819.56	
9	Si	227274.57		
9	Si	212377.56		
2	Sr	89.02	93.37	122.53
2	Sr	109.65		
2	Sr	81.44		
5	Sr	96.32	146.23	
5	Sr	96.39		
5	Sr	245.99		
9	Sr	141.65	128.01	
9	Sr	105.83		
9	Sr	136.55		
2	Ti	1837.99	2067.96	2117.32
2	Ti	2656.74		
2	Ti	1709.14		
5	Ti	1927.96	2338.61	
5	Ti	2103.14		
5	Ti	2984.74		
9	Ti	1892.83	1945.40	
9	Ti	2006.63		
9	Ti	1936.74		
2	Zn	65.25	72.34	78.73
2	Zn	83.14		
2	Zn	68.63		
5	Zn	75.74	85.20	
5	Zn	75.32		
5	Zn	104.54		
9	Zn	78.27	78.67	
9	Zn	79.53		
9	Zn	78.20		
2	Zr	53.98	121.77	140.26
2	Zr	221.98		
2	Zr	89.36		
5	Zr	220.06	197.39	
5	Zr	174.73		
9	Zr	71.37	101.61	
9	Zr	163.33		
9	Zr	70.15		

Element	Container	2 week weathering cycle: totals (aqua regia/HF)		TRT2
		Concentration	Triplicate average	Replicate average
			mg kg-1	
3	Al	73954.36	73796.34	75879.55
3	Al	69039.81		
3	Al	78394.85		
4	Al	68744.92	75172.42	
4	Al	82609.57		
4	Al	74162.77		
7	Al	79696.01	78669.90	
7	Al	76755.52		
7	Al	79558.18		
3	Ba		407.72	407.72
3	Ba			
3	Ba			
4	Ba	392.49		
4	Ba			
4	Ba			
7	Ba	423.73		
7	Ba			
7	Ba	406.94		
3	Ca	11114.57	10623.95	7827.72
3	Ca	10398.69		
3	Ca	10358.60		
4	Ca	1674.28	7713.13	
4	Ca	10410.14		
4	Ca	11054.97		
7	Ca	1981.44	5146.08	
7	Ca	11574.10		
7	Ca	1882.69		
3	Cd	2.00	2.00	2.00
3	Cd	2.00		
3	Cd	2.00		
4	Cd	9.62	2.00	
4	Cd	5.53		
4	Cd	2.00		
7	Cd	2.00	2.00	
7	Cd	2.00		
7	Cd	2.00		
3	Co	10.16	11.44	12.36
3	Co	11.32		
3	Co	12.84		
4	Co	18.94	16.62	
4	Co	17.36		
4	Co	13.57		
7	Co	7.80	9.03	
7	Co	11.72		
7	Co	7.56		
3	Cr	60.98	62.64	65.13
3	Cr	58.79		
3	Cr	68.15		
4	Cr	58.60	63.75	
4	Cr	70.99		
4	Cr	61.66		
7	Cr	72.55	69.02	
7	Cr	65.09		
7	Cr	69.42		
3	Cu	2.00	10.81	13.70
3	Cu	13.68		
3	Cu	16.74		
4	Cu	25.09	17.39	
4	Cu	10.84		
4	Cu	16.26		
7	Cu	5.65	12.92	
7	Cu	26.61		
7	Cu	6.50		
3	Fe	38213.07	37992.51	40047.34
3	Fe	35249.71		
3	Fe	40514.77		
4	Fe	36225.23	39579.99	
4	Fe	44249.33		
4	Fe	38265.40		
7	Fe	44434.17	42569.51	
7	Fe	39965.94		

7	Fe	43308.42		
3	K	8552.86	7319.08	7445.15
3	K	6803.60		
3	K	8600.77		
4	K	5006.42	7044.73	
4	K	8682.33		
4	K	7445.44		
7	K	7405.52	7971.64	
7	K	9563.54		
7	K	6945.86		
3	Mg	10555.38	10383.99	8684.24
3	Mg	10027.79		
3	Mg	10568.81		
4	Mg	4874.14	8558.20	
4	Mg	10227.16		
4	Mg	10573.31		
7	Mg	5262.62	7110.54	
7	Mg	10941.89		
7	Mg	5127.10		
3	Mn	243.30	232.10	214.99
3	Mn	206.46		
3	Mn	246.53		
4	Mn	126.11	224.27	
4	Mn	279.00		
4	Mn	267.69		
7	Mn	137.57	188.60	
7	Mn	291.81		
7	Mn	136.43		
3	Mo	5.00	5.00	5.00
3	Mo	5.00		
3	Mo	5.00		
4	Mo	5.00	5.00	
4	Mo	5.00		
4	Mo	5.00		
7	Mo	5.00	5.00	
7	Mo	5.00		
7	Mo	5.00		
3	Na	25477.83	23864.10	18005.30
3	Na	23221.90		
3	Na	22892.59		
4	Na	4823.55	18130.26	
4	Na	25422.71		
4	Na	24144.52		
7	Na	2237.88	12021.55	
7	Na	25431.83		
7	Na	8394.94		
3	Ni	10.00	10.00	10.00
3	Ni	10.00		
3	Ni	10.00		
4	Ni	10.00	10.00	
4	Ni	10.00		
4	Ni	10.00		
7	Ni	10.00	10.00	
7	Ni	10.00		
7	Ni	17.73		
3	P	848.95	826.42	864.96
3	P	743.63		
3	P	886.68		
4	P	728.66	817.90	
4	P	960.18		
4	P	764.87		
7	P	966.98	950.57	
7	P	890.75		
7	P	993.99		
3	Pb	117.70	121.04	124.16
3	Pb	125.05		
3	Pb	120.37		
4	Pb	130.37	130.97	
4	Pb	116.88		
4	Pb	145.65		
7	Pb	115.20	120.49	
7	Pb	136.24		
7	Pb	110.02		
3	S	4921.46	4153.85	4684.82

3	S	3574.98		
3	S	3965.11		
4	S	3384.54	4496.57	
4	S	6185.59		
4	S	3919.58		
7	S	6142.39	5404.05	
7	S	4134.51		
7	S	5935.24		
3	Si	123681.15	166238.93	174315.33
3	Si	173326.74		
3	Si	201708.89		
4	Si	200271.65	180430.10	
4	Si	141091.48		
4	Si	199927.19		
7	Si	123215.24	176276.96	
7	Si	220772.55		
7	Si	184843.08		
3	Sr	142.50	142.39	126.51
3	Sr	141.65		
3	Sr	143.03		
4	Sr	97.83	119.24	
4	Sr	134.50		
4	Sr	125.39		
7	Sr	92.79	117.90	
7	Sr	170.86		
7	Sr	90.06		
3	Ti	1706.69	1740.94	1803.96
3	Ti	1630.91		
3	Ti	1885.22		
4	Ti	1613.44	1783.79	
4	Ti	2024.57		
4	Ti	1713.35		
7	Ti	1949.00	1887.14	
7	Ti	1835.48		
7	Ti	1876.95		
3	Zn	84.28	76.78	83.50
3	Zn	67.71		
3	Zn	78.35		
4	Zn	75.64	94.42	
4	Zn	101.43		
4	Zn	106.19		
7	Zn	77.84	79.30	
7	Zn	83.81		
7	Zn	76.25		
3	Zr	73.99	61.82	108.17
3	Zr	52.31		
3	Zr	59.17		
4	Zr	145.33	100.66	
4	Zr	99.47		
4	Zr	57.19		
7	Zr	78.07	162.02	
7	Zr	57.64		
7	Zr	350.36		

Element	2 week weathering cycle: totals (aqua regia/HF)			TRTS	
	Container	Concentration	Triplicate average	Replicate average	
			mg kg-1		
6	Al	95563.19	86832.45	80578.17	
6	Al	78101.71			
14	Al	77728.08	76831.51		
14	Al	71947.94			
14	Al	80818.52			
15	Al	80846.76	78070.57		
15	Al	81682.79			
15	Al	71682.16			
6	Ba	554.73	529.26	514.99	
6	Ba	503.79			
14	Ba	491.82	483.87		
14	Ba	462.96			
14	Ba	496.84			
15	Ba	550.31	531.84		
15	Ba	513.36			
15	Ba				
6	Ca	2687.24	2534.77	2693.48	
6	Ca	2382.30			
14	Ca	2884.03	2862.54		
14	Ca	2754.02			
14	Ca	2949.57			
15	Ca	2770.07	2683.15		
15	Ca	2596.23			
15	Ca				
6	Cd	2.00	2.00	2.00	
6	Cd	2.00			
14	Cd	2.00	2.00		
14	Cd	2.00			
14	Cd	9.38			
15	Cd	2.00	2.00		
15	Cd	2.00			
15	Cd	2.44			
6	Co	15.60	15.27	16.81	
6	Co	14.95			
14	Co	13.86	17.26		
14	Co	14.87			
14	Co	23.05			
15	Co	18.64	17.91		
15	Co	16.02			
15	Co	19.08			
6	Cr	85.37	76.76	71.92	
6	Cr	68.15			
14	Cr	66.76	65.45		
14	Cr	62.04			
14	Cr	67.55			
15	Cr	77.69	73.55		
15	Cr	71.77			
15	Cr	71.20			
6	Cu	35.84	40.64	41.17	
6	Cu	45.43			
14	Cu	28.96	37.74		
14	Cu	40.04			
14	Cu	44.22			
15	Cu	47.72	45.15		
15	Cu	42.82			
15	Cu	44.90			
6	Fe	48099.64	43005.42	40232.41	
6	Fe	37911.21			
14	Fe	38352.18	37496.85		
14	Fe	34871.61			
14	Fe	39266.75			
15	Fe	42379.01	40194.96		
15	Fe	39216.75			
15	Fe	38989.12			
6	K	5311.54	5126.89	5347.20	
6	K	4942.23			
14	K	5286.34	5159.48		
14	K	4685.49			
14	K	5506.61			
15	K	5712.45	5755.24		
15	K	4428.75			

15	K	7124.52		
6	Mg	5760.83	5544.28	5989.36
6	Mg	5327.72		
14	Mg	5159.28	5120.51	
14	Mg	4897.43		
14	Mg	5304.81		
15	Mg	5818.49	7303.31	
15	Mg	5445.00		
15	Mg	10646.45		
6	Mn	143.92	134.81	150.61
6	Mn	125.69		
14	Mn	135.14	137.05	
14	Mn	128.01		
14	Mn	148.01		
15	Mn	143.34	179.96	
15	Mn	135.67		
15	Mn	260.88		
6	Mo	5.00	5.00	5.10
6	Mo	5.00		
14	Mo	5.97	5.32	
14	Mo	5.00		
14	Mo	5.01		
15	Mo	5.00	5.00	
15	Mo	5.00		
15	Mo	5.00		
6	Na	4630.34	3859.21	4687.49
6	Na	3088.09		
14	Na	2940.74	4899.57	
14	Na	5738.91		
14	Na	6019.06		
15	Na	2445.80	5303.68	
15	Na	8161.57		
15	Na			
6	Ni	36.98	31.57	23.81
6	Ni	26.16		
14	Ni	19.05	17.74	
14	Ni	16.07		
14	Ni	18.11		
15	Ni	28.00	22.11	
15	Ni	17.90		
15	Ni	20.43		
6	P	1335.96	1166.18	1110.16
6	P	996.41		
14	P	1128.27	1103.20	
14	P	1009.43		
14	P	1171.89		
15	P	1118.13	1061.10	
15	P	1029.20		
15	P	1035.98		
6	Pb	151.43	139.06	133.77
6	Pb	126.89		
14	Pb	121.89	124.96	
14	Pb	129.85		
14	Pb	123.13		
15	Pb	137.52	137.28	
15	Pb	129.68		
15	Pb	144.66		
6	S	5718.65	4431.34	3856.02
6	S	3144.04		
14	S	4194.52	3729.57	
14	S	3036.04		
14	S	3958.16		
15	S	3555.49	3407.15	
15	S	3208.47		
15	S	3457.50		
6	Si	169514.34	157041.56	166231.53
6	Si	144568.79		
14	Si	115013.30	142848.27	
14	Si	182837.02		
14	Si	130694.49		
15	Si	184526.06	198804.77	
15	Si	229504.14		
15	Si	182384.11		
6	Sr	200.78	192.92	194.48

6	Sr	185.06		
14	Sr	173.69	176.85	
14	Sr	167.01		
14	Sr	189.86		
15	Sr	202.14	213.67	
15	Sr	182.22		
15	Sr	246.65		
6	Ti	2965.76	2670.47	2481.78
6	Ti	2375.18		
14	Ti	2325.45	2268.85	
14	Ti	2111.92		
14	Ti	2369.20		
15	Ti	2602.51	2506.03	
15	Ti	2450.27		
15	Ti	2465.32		
6	Zn	96.61	89.77	88.99
6	Zn	82.94		
14	Zn	84.57	86.43	
14	Zn	78.09		
14	Zn	96.62		
15	Zn	92.66	90.77	
15	Zn	85.13		
15	Zn	94.53		
6	Zr	195.36	151.51	139.15
6	Zr	107.66		
14	Zr	95.26	146.14	
14	Zr	154.94		
14	Zr	188.23		
15	Zr	72.59	119.81	
15	Zr	213.65		
15	Zr	73.20		

Element	Container	2 week weathering cycle: totals (aqua regia/HF)		TRT4
		Concentration	Triplicate average	Replicate average
			mg kg ⁻¹	
8	Al	77461.43	77369.51	85649.18
8	Al	74393.62		
8	Al	80253.48		
10	Al	91235.98	93077.16	
10	Al	94918.35		
13	Al	86664.30	86500.88	
13	Al	96935.41		
13	Al	75902.94		
8	Ba	497.86	516.21	542.67
8	Ba			
8	Ba	534.57		
10	Ba	570.07	557.70	
10	Ba	545.33		
13	Ba		554.10	
13	Ba	609.20		
13	Ba	499.00		
8	Ca	2977.12	3033.72	3095.65
8	Ca			
8	Ca	3090.32		
10	Ca	3247.71	3174.52	
10	Ca	3101.32		
13	Ca		3078.72	
13	Ca	3430.30		
13	Ca	2727.14		
8	Cd	2.00	2.00	2.00
8	Cd	2.00		
8	Cd	2.00		
10	Cd	2.00	2.00	
10	Cd	2.00		
13	Cd	2.00	2.00	
13	Cd	2.00		
13	Cd	2.00		
8	Co	17.91	18.09	17.73
8	Co	16.87		
8	Co	19.48		
10	Co	18.08	17.13	
10	Co	16.17		
13	Co	17.81	17.99	
13	Co	19.14		
13	Co	17.01		
8	Cr	67.19	65.93	73.91
8	Cr	60.35		
8	Cr	70.25		
10	Cr	79.59	80.91	
10	Cr	82.24		
13	Cr	74.43	74.89	
13	Cr	84.59		
13	Cr	65.66		
8	Cu	53.72	47.74	44.04
8	Cu	29.62		
8	Cu	59.90		
10	Cu	42.55	40.23	
10	Cu	37.91		
13	Cu	35.23	44.15	
13	Cu	47.38		
13	Cu	49.83		
8	Fe	35674.84	35314.49	40255.07
8	Fe	33997.67		
8	Fe	36270.96		
10	Fe	43603.41	44456.15	
10	Fe	45308.90		
13	Fe	41216.72	40994.57	
13	Fe	45985.08		
13	Fe	35781.90		
8	K	6225.93	6078.29	6655.78
8	K	6938.33		
8	K	5070.60		
10	K	7288.72	6596.57	
10	K	5924.41		
13	K	8529.11	7292.50	
13	K	7764.32		

13	K	5584.08		
8	Mg	5077.64	7421.55	6840.04
8	Mg	11961.04		
8	Mg	5225.97		
10	Mg	5881.59	5713.25	
10	Mg	5544.91		
13	Mg	10886.17	7385.32	
13	Mg	6226.89		
13	Mg	5042.91		
8	Mn	128.15	175.90	44.04
8	Mn	276.12		
8	Mn	123.44		
10	Mn	139.43	136.42	
10	Mn	133.41		
13	Mn	248.48	175.01	
13	Mn	154.81		
13	Mn	121.76		
8	Mo	5.00	5.00	5.00
8	Mo	5.01		
8	Mo	5.46		
10	Mo	5.00	5.00	
10	Mo	5.00		
13	Mo	5.00	5.00	
13	Mo	5.00		
8	Na	8419.67	6527.71	7058.63
8	Na			
8	Na	4635.76		
10	Na	5316.62	8495.65	
10	Na	11674.68		
13	Na		6152.53	
13	Na	8307.50		
13	Na	3997.55		
8	Ni	18.82	21.17	27.97
8	Ni	17.17		
8	Ni	27.54		
10	Ni	33.01	32.63	
10	Ni	32.26		
13	Ni	27.35	30.11	
13	Ni	36.60		
13	Ni	26.39		
8	P	1087.32	1087.12	1215.77
8	P	1077.69		
8	P	1096.37		
10	P	1307.26	1336.70	
10	P	1366.15		
13	P	1263.95	1223.49	
13	P	1359.38		
13	P	1047.15		
8	Pb	138.90	130.88	139.09
8	Pb	109.20		
8	Pb	144.56		
10	Pb	160.99	144.39	
10	Pb	127.78		
13	Pb	125.04	142.01	
13	Pb	172.59		
13	Pb	128.41		
8	S	2888.36	3282.71	4066.14
8	S	4214.68		
8	S	2745.10		
10	S	4494.75	4734.87	
10	S	4974.98		
13	S	4932.51	4180.84	
13	S	4823.61		
13	S	2786.40		
8	Si	223453.81	181324.61	177843.81
8	Si	156560.99		
8	Si	163959.03		
10	Si	155525.16	191368.50	
10	Si	227211.84		
13	Si	175808.58	160838.32	
13	Si	177853.12		
13	Si	128853.26		
8	Sr	214.41	233.33	235.90

8	Sr	242.23		
8	Sr	243.34		
10	Sr	240.04	236.39	
10	Sr	232.74		
13	Sr	250.48	237.98	
13	Sr	256.21		
13	Sr	207.24		
13	Sr	2453.85	2502.12	2783.15
8	Ti	2376.97		
8	Ti	2675.54		
8	Ti	2980.71	3015.60	
10	Ti	3050.49		
13	Ti	2895.91	2831.75	
13	Ti	3173.99		
13	Ti	2425.36		
13	Ti	87.63	96.97	95.48
8	Zn	113.48		
8	Zn	89.80		
8	Zn	99.89	97.83	
10	Zn	95.76		
10	Zn	88.38	91.63	
13	Zn	104.66		
13	Zn	81.87		
13	Zn	256.75	195.04	209.42
8	Zr	219.75		
8	Zr	108.63		
8	Zr	234.84	260.12	
10	Zr	285.40		
10	Zr	115.51	173.11	
13	Zr	257.26		
13	Zr	146.55		

Element	Container	2 week weathering cycle: totals (aqua regia/HF)		TRT5 Replicate average
		Concentration	Triplicate average mg kg-1	
1	Al	85151.72	86541.22	86743.94
1	Al	90568.53		
1	Al	83903.43		
11	Al	85094.85	82817.50	
11	Al	80540.15		
12	Al	99502.38	90873.11	
12	Al	82122.26		
12	Al	90994.70		
1	Ba	513.76	513.76	556.49
1	Ba			
1	Ba			
11	Ba		564.04	
11	Ba	559.31		
11	Ba	568.78		
12	Ba	661.80	591.69	
12	Ba	577.68		
12	Ba	535.59		
1	Ca	3258.96	10179.78	7979.11
1	Ca	13957.54		
1	Ca	13322.85		
11	Ca	11012.03	5778.45	
11	Ca	3223.08		
11	Ca	3100.23		
12	Ca	3869.53	3346.63	
12	Ca	3215.08		
12	Ca	2955.27		
1	Cd	2.00	2.00	2.00
1	Cd	2.00		
1	Cd	2.00		
11	Cd	2.00	2.00	
11	Cd	2.00		
11	Cd	2.00		
12	Cd	2.00	2.00	
12	Cd	2.00		
12	Cd	2.00		
1	Co	16.85	20.31	20.42
1	Co	24.21		
1	Co	19.89		
11	Co	22.88	20.54	
11	Co	19.23		
11	Co	19.51		
12	Co	24.09	20.48	
12	Co	20.51		
12	Co	16.83		
1	Cr	75.21	74.48	72.21
1	Cr	75.68		
1	Cr	72.55		
11	Cr	68.34	69.94	
11	Cr	71.61		
11	Cr	69.86		
12	Cr	88.34	79.90	
12	Cr	71.77		
12	Cr	79.59		
1	Cu	37.51	41.20	49.20
1	Cu	50.04		
1	Cu	36.06		
11	Cu	63.51	57.20	
11	Cu	47.20		
11	Cu	60.91		
12	Cu	59.05	53.67	
12	Cu	60.91		
12	Cu	41.05		
1	Fe	41075.43	40761.65	39233.70
1	Fe	41674.55		
1	Fe	39534.96		
11	Fe	36654.26	37705.76	
11	Fe	39323.36		
11	Fe	37139.68		
12	Fe	46639.28	42712.96	
12	Fe	37527.39		
12	Fe	43972.20		

1	K	6230.85	8413.11	6843.40
1	K	11398.68		
1	K	7609.81		
11	K	6611.04	5273.69	
11	K	4460.35		
11	K	4749.68		
12	K	8920.70	8687.44	
12	K	8408.22		
12	K	8733.40		
1	Mg	5026.33	9153.99	7990.27
1	Mg	11734.15		
1	Mg	10701.50		
11	Mg	9813.26	6826.55	
11	Mg	5387.70		
11	Mg	5278.69		
12	Mg	6362.49	5733.53	
12	Mg	5404.14		
12	Mg	5433.96		
1	Mn	123.46	235.56	201.54
1	Mn	309.27		
1	Mn	273.94		
11	Mn	245.13	167.52	
11	Mn	130.00		
11	Mn	127.45		
12	Mn	155.15	137.48	
12	Mn	127.73		
12	Mn	129.55		
1	Mo	5.00	5.00	5.00
1	Mo	5.00		
1	Mo	5.00		
11	Mo	5.00	5.00	
11	Mo	5.00		
11	Mo	5.00		
12	Mo	5.00	5.00	
12	Mo	5.00		
12	Mo	5.00		
1	Na	6748.03	6748.03	4982.11
1	Na			
1	Na			
11	Na		4109.32	
11	Na	5466.22		
11	Na	2752.43		
12	Na	3511.81	4088.97	
12	Na	2980.13		
12	Na	5774.97		
1	Ni	31.69	32.07	29.23
1	Ni	34.52		
1	Ni	30.00		
11	Ni	19.97	26.39	
11	Ni	30.75		
11	Ni	28.46		
12	Ni	43.01	38.55	
12	Ni	36.03		
12	Ni	36.60		
1	P	1308.86	1306.97	1243.46
1	P	1325.68		
1	P	1286.39		
11	P	1136.77	1179.95	
11	P	1269.09		
11	P	1133.99		
12	P	1497.63	1317.21	
12	P	1176.05		
12	P	1277.94		
1	Pb	125.76	161.90	148.23
1	Pb	187.78		
1	Pb	172.17		
11	Pb	144.85	134.56	
11	Pb	117.25		
11	Pb	141.59		
12	Pb	183.19	156.04	
12	Pb	156.72		
12	Pb	128.21		
1	S	4586.62	4600.29	3871.18
1	S	4349.48		

1	S	4864.78		
11	S	2963.97	3142.07	
11	S	3743.17		
11	S	2719.08		
12	S	4531.30	3992.92	
12	S	2767.88		
12	S	4679.59		
1	Si	143992.66	160801.91	156994.92
1	Si	176755.74		
1	Si	161657.33		
11	Si	174538.16	153187.94	
11	Si	136192.31		
11	Si	148833.34		
12	Si	152286.25	156000.22	
12	Si	157147.95		
12	Si	158566.47		
1	Sr	231.57	278.89	269.51
1	Sr	310.46		
1	Sr	294.63		
11	Sr	275.59	260.14	
11	Sr	253.50		
11	Sr	251.34		
12	Sr	295.65	256.39	
12	Sr	251.20		
12	Sr	222.32		
1	Ti	2954.69	3005.16	2934.21
1	Ti	3140.35		
1	Ti	2920.46		
11	Ti	2785.75	2863.26	
11	Ti	2987.09		
11	Ti	2816.94		
12	Ti	3477.07	3073.07	
12	Ti	2779.48		
12	Ti	2962.68		
1	Zn	92.06	97.98	98.72
1	Zn	111.28		
1	Zn	90.59		
11	Zn	103.86	99.46	
11	Zn	97.48		
11	Zn	97.05		
12	Zn	113.95	100.61	
12	Zn	93.30		
12	Zn	94.57		
1	Zr	187.21	141.49	123.80
1	Zr	131.89		
1	Zr	105.58		
11	Zr	92.26	106.10	
11	Zr	139.58		
11	Zr	86.47		
12	Zr	102.02	106.97	
12	Zr	132.58		
12	Zr	86.31		

18 week weathering cycle: totals (aqua regia/HF)			TRT1	
Element	Container	Concentration	Triplicate average	Replicate average
			mg kg ⁻¹	
2	Al	75032.94	79652.45	78749.04
2	Al	80620.16		
2	Al	83304.26		
5	Al	78835.30	75982.51	
5	Al	71997.93		
5	Al	77114.30		
9	Al	78331.93	80612.16	
9	Al	83213.55		
9	Al	80291.01		
2	Ba	495.35	457.34	457.34
2	Ba			
2	Ba			
5	Ba	411.19		
5	Ba			
5	Ba			
9	Ba	465.47		
9	Ba			
9	Ba			
2	Ca	26391.36	27157.96	28079.82
2	Ca			
2	Ca	27923.36		
5	Ca	35514.73	30475.74	
5	Ca			
5	Ca	25436.76		
9	Ca	26489.15	26606.35	
9	Ca	26723.54		
9	Ca			
2	Cd	2.00	2.00	2.51
2	Cd	2.00		
2	Cd	2.00		
5	Cd	2.00	3.52	
5	Cd	6.26		
5	Cd	2.29		
9	Cd	2.00	2.00	
9	Cd	2.00		
9	Cd	2.00		
2	Co	15.26	14.83	15.24
2	Co	10.78		
2	Co	18.45		
5	Co	17.99	16.81	
5	Co	15.77		
5	Co	16.66		
9	Co	15.83	14.10	
9	Co	17.03		
9	Co	9.44		
2	Cr	63.37	69.29	68.42
2	Cr	73.87		
2	Cr	70.63		
5	Cr	65.47	64.71	
5	Cr	64.33		
5	Cr	64.33		
9	Cr	67.00	71.26	
9	Cr	74.64		
9	Cr	72.15		
2	Cu	11.79	12.99	15.55
2	Cu	11.79		
2	Cu	15.40		
5	Cu	17.74	21.39	
5	Cu	28.68		
5	Cu	17.75		
9	Cu	11.77	12.28	
9	Cu	12.52		
9	Cu	12.55		
2	Fe	37132.93	40998.37	40611.98
2	Fe	43921.81		
2	Fe	41940.36		
5	Fe	38686.64	38610.24	
5	Fe	38331.41		
5	Fe	38812.67		
9	Fe	40493.20	42227.34	
9	Fe	43293.46		

9	Fe	42895.36		
2	K	11681.64	11061.19	11082.58
2	K	9370.99		
2	K	12130.94		
5	K	12901.16	10547.71	
5	K	7638.00		
5	K	11103.98		
9	K	13543.00	11638.85	
9	K	14441.59		
9	K	6931.96		
2	Mg	19812.85	20594.15	21214.18
2	Mg			
2	Mg	21375.45		
5	Mg	25127.90	22390.08	
5	Mg			
5	Mg	19652.26		
9	Mg	20403.32	20658.30	
9	Mg	20913.29		
9	Mg			
2	Mn	513.87	546.81	522.22
2	Mn			
2	Mn	579.75		
5	Mn	634.30	529.31	
5	Mn			
5	Mn	424.32		
9	Mn	435.57	490.54	
9	Mn	545.51		
9	Mn			
2	Mo	5.00	5.00	5.00
2	Mo	5.00		
2	Mo	5.00		
5	Mo	5.00	5.00	
5	Mo	5.00		
5	Mo	5.00		
9	Mo	5.00	5.00	
9	Mo	5.00		
9	Mo	5.00		
2	Na	66212.47	68735.60	71334.49
2	Na			
2	Na	71258.74		
5	Na	89042.92	76721.63	
5	Na			
5	Na	64400.34		
9	Na	67799.61	68546.23	
9	Na	69292.85		
9	Na			
2	Ni	10.00	12.02	11.23
2	Ni	16.07		
2	Ni	10.00		
5	Ni	10.00	10.00	
5	Ni	10.00		
5	Ni	10.00		
9	Ni	11.48	11.67	
9	Ni	13.54		
9	Ni	10.00		
2	P	748.52	766.32	750.24
2	P	818.71		
2	P	731.72		
5	P	686.45	714.04	
5	P	713.69		
5	P	741.98		
9	P	806.33	770.37	
9	P	751.30		
9	P	753.48		
2	Pb	184.40	166.00	155.07
2	Pb	161.34		
2	Pb	152.27		
5	Pb	146.74	137.79	
5	Pb	123.09		
5	Pb	143.55		
9	Pb	144.32	161.42	
9	Pb	217.15		
9	Pb	122.79		
2	S	3632.20	3707.85	3637.16

2	S	3343.48		
2	S	4147.88		
5	S	4161.10	3598.46	
5	S	2978.07		
5	S	3656.21		
9	S	3689.04	3605.18	
9	S	3886.29		
9	S	3240.23		
2	Si	301961.49	281782.26	296405.30
2	Si	205528.55		
2	Si	337856.75		
5	Si	350000.00	320684.34	
5	Si	324313.90		
5	Si	287739.11		
9	Si	302674.07	286749.31	
9	Si	305840.33		
9	Si	251733.54		
2	Sr	288.13	197.09	189.10
2	Sr	101.28		
2	Sr	201.87		
5	Sr	225.16	184.05	
5	Sr	101.97		
5	Sr	225.02		
9	Sr	209.45	186.16	
9	Sr	242.52		
9	Sr	106.51		
2	Ti	2006.27	2030.52	2052.58
2	Ti	1990.95		
2	Ti	2094.35		
5	Ti	2090.68	1979.27	
5	Ti	1825.45		
5	Ti	2021.68		
9	Ti	2101.08	2147.94	
9	Ti	2249.91		
9	Ti	2092.83		
2	Zn	86.68	82.55	83.77
2	Zn	78.41		
2	Zn			
5	Zn	116.43	87.54	
5	Zn	68.83		
5	Zn	77.36		
9	Zn	72.21	81.23	
9	Zn	102.73		
9	Zn	68.76		
2	Zr	68.32	101.21	146.63
2	Zr	110.87		
2	Zr	124.44		
5	Zr	143.20	190.57	
5	Zr	352.58		
5	Zr	75.94		
9	Zr	91.65	148.13	
9	Zr	126.27		
9	Zr	226.46		

18 week weathering cycle: totals (aqua regia/HF)				
Element	Container	Concentration	Triplicate average	TRT2 Replicate average
			mg kg-1	
3	Al	75870.48	76129.08	75513.49
3	Al	78882.74		
3	Al	73634.03		
4	Al	71002.82	73745.37	
4	Al	78428.35		
4	Al	71804.95		
7	Al	83511.16	76666.01	
7	Al	73025.28		
7	Al	73461.59		
3	Ba	464.33	434.99	427.96
3	Ba	437.42		
3	Ba	403.21		
4	Ba		415.30	
4	Ba			
4	Ba	415.30		
7	Ba	456.81	433.66	
7	Ba			
7	Ba	410.51		
3	Ca	3269.21	3202.42	13004.70
3	Ca	3286.00		
3	Ca	3052.07		
4	Ca	27567.61	23331.80	
4	Ca	39611.68		
4	Ca	2816.12		
7	Ca	3396.17	12479.88	
7	Ca	30970.29		
7	Ca	3073.17		
3	Cd	2.00	2.00	2.10
3	Cd	2.00		
3	Cd	2.00		
4	Cd	2.00	2.00	
4	Cd	2.00		
4	Cd	2.00		
7	Cd	2.00	2.30	
7	Cd	2.00		
7	Cd	2.90		
3	Co	10.46	10.98	13.34
3	Co	11.79		
3	Co	10.69		
4	Co	16.19	14.59	
4	Co	16.67		
4	Co	10.92		
7	Co	11.24	14.46	
7	Co	18.05		
7	Co	14.10		
3	Cr	66.34	69.16	67.53
3	Cr	72.54		
3	Cr	66.62		
4	Cr	58.98	66.17	
4	Cr	75.97		
4	Cr	63.56		
7	Cr	73.87	67.25	
7	Cr	61.66		
7	Cr	66.24		
3	Cu	71.53	70.30	42.59
3	Cu	71.78		
3	Cu	67.61		
4	Cu	30.00	36.21	
4	Cu	36.17		
4	Cu	42.48		
7	Cu	20.64	21.26	
7	Cu	17.03		
7	Cu	26.12		
3	Fe	42897.70	42679.15	41711.66
3	Fe	44230.53		
3	Fe	40909.24		
4	Fe	35720.06	40532.43	
4	Fe	46569.28		
4	Fe	39307.97		
7	Fe	45723.43	41923.39	
7	Fe	38180.43		

7	Fe	41866.31		
3	K	6033.38	6225.93	8272.71
3	K	5648.27		
3	K	6996.15		
4	K	10783.06	9563.54	
4	K	11938.38		
4	K	5969.19		
7	K	8985.88	9028.67	
7	K	11424.90		
7	K	6675.22		
3	Mg	5376.76	5345.97	10586.66
3	Mg	5524.27		
3	Mg	5136.89		
4	Mg	19179.47	15713.20	
4	Mg	22931.92		
4	Mg	5028.20		
7	Mg	5659.53	10700.80	
7	Mg	21269.21		
7	Mg	5173.67		
3	Mn	177.85	178.74	169.56
3	Mn	183.62		
3	Mn	174.76		
4	Mn		151.56	
4	Mn			
4	Mn	151.56		
7	Mn	184.88	178.38	
7	Mn			
7	Mn	171.88		
3	Mo	5.00	5.33	5.11
3	Mo	5.00		
3	Mo	6.01		
4	Mo	5.00	5.00	
4	Mo	5.00		
4	Mo	5.00		
7	Mo	5.00	5.00	
7	Mo	5.00		
7	Mo	5.00		
3	Na	10851.86	11516.68	31101.23
3	Na	10935.32		
3	Na	12762.86		
4	Na	63431.47	47859.78	
4	Na	69061.51		
4	Na	11086.36		
7	Na	18225.53	33927.24	
7	Na	72023.95		
7	Na	11532.25		
3	Ni	15.38	12.85	11.33
3	Ni	12.39		
3	Ni	10.79		
4	Ni	10.00	10.11	
4	Ni	10.00		
4	Ni	10.33		
7	Ni	13.08	11.02	
7	Ni	10.00		
7	Ni	10.00		
3	P	1077.92	1045.99	1011.53
3	P	1031.06		
3	P	1029.00		
4	P	865.01	967.78	
4	P	1056.58		
4	P	981.77		
7	P	1012.07	1020.81	
7	P	923.45		
7	P	1126.92		
3	Pb	127.47	122.25	129.30
3	Pb	121.87		
3	Pb	117.41		
4	Pb	136.33	135.45	
4	Pb	150.59		
4	Pb	119.44		
7	Pb	139.47	130.19	
7	Pb	136.36		
7	Pb	114.74		
3	S	2824.73	2699.85	2936.25

3	S	2747.50		
3	S	2527.33		
4	S	3106.07	3142.91	
4	S	3707.49		
4	S	2815.16		
7	S	2812.13	2966.00	
7	S	3392.59		
7	S	2693.30		
3	Si	286971.28	275770.40	292776.25
3	Si	281517.58		
3	Si	258822.36		
4	Si	259163.88	287229.53	
4	Si	350000.00		
4	Si	252524.72		
7	Si	350000.00	315328.82	
7	Si	319549.29		
7	Si	276437.16		
3	Sr	92.18	96.13	144.56
3	Sr	99.76		
3	Sr	96.46		
4	Sr	245.82	208.20	
4	Sr	290.61		
4	Sr	88.19		
7	Sr	102.66	129.34	
7	Sr	189.47		
7	Sr	95.90		
3	Ti	1569.98	1646.38	1720.28
3	Ti	1635.75		
3	Ti	1733.42		
4	Ti	1704.03	1766.72	
4	Ti	2016.58		
4	Ti	1579.57		
7	Ti	1842.29	1747.73	
7	Ti	1727.87		
7	Ti	1673.03		
3	Zn	100.18	100.19	102.65
3	Zn	105.22		
3	Zn	95.18		
4	Zn	99.93	104.26	
4	Zn	127.23		
4	Zn	85.61		
7	Zn	100.02	103.50	
7	Zn	118.17		
7	Zn	92.30		
3	Zr	272.52	280.45	226.53
3	Zr	343.89		
3	Zr	224.94		
4	Zr	63.90	90.84	
4	Zr	74.27		
4	Zr	134.35		
7	Zr	118.19	308.30	
7	Zr	692.04		
7	Zr	114.68		

Element	Container	18 week weathering cycle: totals (aqua regia/HF)		TR13
		Concentration	Triplicate average	Replicate average
			mg kg-1	
6	Al	71016.24	79852.14	77437.97
6	Al	82165.63		
6	Al	86374.55		
14	Al	79774.77	76891.26	
14	Al	72617.13		
14	Al	78281.88		
15	Al	77397.16	75570.52	
15	Al	75593.71		
15	Al	73720.69		
6	Ba	458.63	528.64	492.91
6	Ba	523.86		
6	Ba	603.45		
14	Ba		506.29	
14	Ba	512.22		
14	Ba	500.36		
15	Ba		443.81	
15	Ba	444.49		
15	Ba	443.12		
6	Ca	3458.46	3730.61	3496.06
6	Ca	3751.65		
6	Ca	3981.73		
14	Ca		3283.29	
14	Ca	3201.32		
14	Ca	3365.26		
15	Ca		3474.29	
15	Ca	3510.20		
15	Ca	3438.37		
6	Cd	2.00	2.00	2.00
6	Cd	2.00		
6	Cd	2.00		
14	Cd	2.00	2.00	
14	Cd	2.00		
14	Cd	2.00		
15	Cd	2.00	2.00	
15	Cd	2.00		
6	Co	15.07	16.14	17.24
6	Co	15.58		
6	Co	17.78		
14	Co	22.00	16.80	
14	Co	14.05		
14	Co	14.35		
15	Co	19.75	18.78	
15	Co	12.90		
15	Co	23.68		
6	Cr	63.95	72.66	68.23
6	Cr	75.02		
6	Cr	79.03		
14	Cr	67.38	67.06	
14	Cr	65.09		
14	Cr	68.72		
15	Cr	64.71	64.96	
15	Cr	66.43		
15	Cr	63.76		
6	Cu	43.70	55.58	46.60
6	Cu	52.97		
6	Cu	70.07		
14	Cu	48.83	44.53	
14	Cu	40.83		
14	Cu	43.92		
15	Cu	37.17	39.71	
15	Cu	37.73		
15	Cu	44.22		
6	Fe	37462.89	41733.17	39371.40
6	Fe	42838.71		
6	Fe	44898.11		
14	Fe	38362.85	38418.55	
14	Fe	37437.48		
14	Fe	39455.31		
15	Fe	37886.01	37962.49	
15	Fe	38372.21		

15	Fe	37629.26		
6	K	5006.42	6611.04	7366.99
6	K	6803.59		
6	K	8023.11		
14	K	12195.12	7830.55	
14	K	5327.34		
14	K	5969.19		
15	K	11296.53	7659.39	
15	K	6354.30		
15	K	5327.34		
6	Mg	4868.42	5445.68	5176.96
6	Mg	5591.70		
6	Mg	5876.92		
14	Mg		5114.62	
14	Mg	5015.94		
14	Mg	5213.30		
15	Mg		4970.58	
15	Mg	5040.86		
15	Mg	4900.29		
6	Mn	155.08	172.60	159.98
6	Mn	177.85		
6	Mn	184.88		
14	Mn		146.78	
14	Mn	144.88		
14	Mn	148.88		
15	Mn		160.56	
15	Mn	156.20		
15	Mn	164.92		
6	Mo	5.00	5.00	5.00
6	Mo	5.00		
6	Mo	5.00		
14	Mo	5.00	5.00	
14	Mo	5.00		
14	Mo	5.00		
15	Mo	5.00	5.00	
15	Mo	5.00		
15	Mo	5.00		
6	Na	6324.50	10019.20	11018.00
6	Na	14027.94		
6	Na	9705.16		
14	Na		10741.63	
14	Na	6626.60		
14	Na	14856.66		
15	Na		12293.16	
15	Na	13119.39		
15	Na	11466.93		
6	Ni	15.38	20.65	17.72
6	Ni	21.11		
6	Ni	25.48		
14	Ni	20.43	16.98	
14	Ni	18.59		
14	Ni	11.93		
15	Ni	12.39	15.53	
15	Ni	20.88		
15	Ni	13.31		
6	P	1156.37	1196.71	1136.71
6	P	1170.48		
6	P	1263.29		
14	P	1038.78	1052.14	
14	P	1068.80		
14	P	1048.83		
15	P	1134.42	1161.27	
15	P	1162.09		
15	P	1187.30		
6	Pb	131.37	152.37	150.91
6	Pb	151.51		
6	Pb	174.23		
14	Pb	218.49	164.73	
14	Pb	139.30		
14	Pb	136.40		
15	Pb	151.81	135.63	
15	Pb	131.35		
15	Pb	123.73		
6	S	2314.99	2483.69	2538.52

6	S	2472.47		
6	S	2663.62		
14	S	3008.79	2600.55	
14	S	2385.30		
14	S	2407.56		
15	S	2902.92	2531.31	
15	S	2365.71		
15	S	2325.29		
6	Si	164161.16	249664.67	260038.21
6	Si	312067.44		
6	Si	272765.41		
14	Si	291148.99	248966.84	
14	Si	179407.30		
14	Si	276344.24		
15	Si	301061.98	281483.12	
15	Si	278905.06		
15	Si	264482.33		
6	Sr	145.37	166.50	190.78
6	Sr	173.76		
6	Sr	180.37		
14	Sr	295.16	203.25	
14	Sr	150.47		
14	Sr	164.11		
15	Sr	295.71	202.60	
15	Sr	151.57		
15	Sr	160.53		
6	Ti	1980.38	2213.56	2130.54
6	Ti	2271.77		
6	Ti	2388.53		
14	Ti	2258.69	2140.32	
14	Ti	2023.48		
14	Ti	2138.80		
15	Ti	2179.03	2037.75	
15	Ti	2002.33		
15	Ti	1931.90		
6	Zn	91.02	108.80	102.93
6	Zn	104.04		
6	Zn	131.34		
14	Zn	123.60	102.32	
14	Zn	88.28		
14	Zn	95.08		
15	Zn	101.45	97.68	
15	Zn	92.73		
15	Zn	98.87		
6	Zr	136.33	208.36	177.54
6	Zr	152.04		
6	Zr	336.72		
14	Zr	128.10	170.54	
14	Zr	123.22		
14	Zr	260.32		
15	Zr	121.24	153.72	
15	Zr	155.24		
15	Zr	184.68		

Element	Container	18 week weathering cycle: totals (aqua regia/HF)		
		Concentration	Triplicate average	
			mg kg ⁻¹	TRT4
				Replicate average
8	Al	84213.23	80358.29	83210.62
8	Al	78738.00		
8	Al	78123.66		
10	Al	86781.17	84297.11	
10	Al	89675.90		
10	Al	76434.26		
13	Al	81867.03	84976.46	
13	Al	86182.00		
13	Al	86880.36		
8	Ba	527.73	508.92	525.97
8	Ba			
8	Ba	490.10		
10	Ba	564.68	529.86	
10	Ba	541.42		
10	Ba	483.49		
13	Ba	535.26	539.13	
13	Ba			
13	Ba	543.01		
8	Ca	3879.62	3756.14	3903.57
8	Ca			
8	Ca	3632.67		
10	Ca	3997.41	4102.88	
10	Ca	4572.60		
10	Ca	3738.62		
13	Ca	3658.36	3851.69	
13	Ca			
13	Ca	4045.02		
8	Cd	2.00	2.00	2.00
8	Cd	2.00		
8	Cd	2.00		
10	Cd		2.00	
10	Cd	2.00		
10	Cd	2.00		
13	Cd	2.00	2.00	
13	Cd	2.00		
13	Cd	2.00		
8	Co	18.92	19.64	19.53
8	Co	23.19		
8	Co	16.82		
10	Co	24.01	19.72	
10	Co	18.14		
10	Co	17.00		
13	Co	17.12	19.23	
13	Co	23.07		
13	Co	17.49		
8	Cr	75.59	70.43	73.09
8	Cr	66.24		
8	Cr	69.48		
10	Cr	77.88	75.72	
10	Cr	79.03		
10	Cr	70.25		
13	Cr	70.82	73.11	
13	Cr	72.92		
13	Cr	75.59		
8	Cu	54.73	51.46	51.17
8	Cu	50.35		
8	Cu	49.32		
10	Cu	62.23	58.61	
10	Cu	60.65		
10	Cu	52.94		
13	Cu	42.07	43.43	
13	Cu	43.07		
13	Cu	45.16		
8	Fe	40678.47	38556.44	40128.17
8	Fe	36516.53		
8	Fe	38474.33		
10	Fe	41449.75	40425.89	
10	Fe	41795.88		
10	Fe	38032.04		
13	Fe	41500.94	41402.19	
13	Fe	40509.05		

13	Fe	42196.59		
8	K	5840.82	7937.53	7737.84
8	K	11681.64		
8	K	6290.12		
10	K	6482.67	7231.49	
10	K	8985.88		
10	K	6225.93		
13	K	5648.27	8044.50	
13	K	10718.87		
13	K	7766.37		
8	Mg	5523.46	5357.14	8721.98
8	Mg			
8	Mg	5190.83		
10	Mg	5688.54	5511.06	
10	Mg	5666.48		
10	Mg	5178.16		
13	Mg	5306.88	5438.05	
13	Mg			
13	Mg	5569.22		
8	Mn	166.04	157.71	159.57
8	Mn			
8	Mn	149.38		
10	Mn	164.22	154.75	
10	Mn	156.90		
10	Mn	143.13		
13	Mn	162.88	166.25	
13	Mn			
13	Mn	169.63		
8	Mo	5.00	5.00	5.00
8	Mo	5.00		
8	Mo	5.00		
10	Mo	5.00	5.00	
10	Mo	5.00		
10	Mo	5.00		
13	Mo	5.00	5.00	
13	Mo	5.00		
13	Mo	5.00		
8	Na	16801.24	13024.59	13450.74
8	Na			
8	Na	9247.94		
10	Na	13301.73	12401.34	
10	Na	19640.75		
10	Na	4261.54		
13	Na	9922.89	14926.29	
13	Na			
13	Na	19929.69		
8	Ni	17.44	17.29	23.05
8	Ni	16.29		
8	Ni	18.13		
10	Ni	28.46	25.24	
10	Ni	25.02		
10	Ni	22.26		
13	Ni	22.49	26.62	
13	Ni	28.46		
13	Ni	28.92		
8	P	1170.35	1142.98	1184.36
8	P	1072.63		
8	P	1185.98		
10	P	1205.33	1192.40	
10	P	1213.20		
10	P	1158.69		
13	P	1283.78	1217.70	
13	P	1172.13		
13	P	1197.20		
8	Pb	138.23	165.70	152.06
8	Pb	228.48		
8	Pb	130.39		
10	Pb	158.47	144.48	
10	Pb	146.67		
10	Pb	128.29		
13	Pb	141.71	146.00	
13	Pb	152.19		
13	Pb	144.10		
8	S	2342.44	2538.57	2522.65

8	S	2861.89		
8	S	2411.39		
10	S	2351.98	2360.48	
10	S	2444.81		
10	S	2284.66		
13	S	2491.06	2668.90	
13	S	3155.27		
13	S	2360.39		
8	Si	316017.52	283556.71	286704.59
8	Si	279257.54		
8	Si	255395.07		
10	Si	293784.15	258973.93	
10	Si	350000.00		
10	Si	133137.63		
13	Si	285597.26	317583.13	
13	Si	317152.13		
13	Si	350000.00		
8	Sr	210.83	234.57	225.45
8	Sr	303.70		
8	Sr	189.19		
10	Sr	219.23	212.75	
10	Sr	227.36		
10	Sr	191.67		
13	Sr	186.71	229.01	
13	Sr	292.95		
13	Sr	207.38		
8	Ti	2576.97	2463.26	2502.53
8	Ti	2419.98		
8	Ti	2392.83		
10	Ti	2658.96	2621.89	
10	Ti	2778.67		
10	Ti	2428.05		
13	Ti	2347.94	2422.43	
13	Ti	2500.99		
13	Ti	2418.37		
8	Zn	101.10	106.47	116.44
8	Zn	122.23		
8	Zn	96.09		
10	Zn	111.49	104.46	
10	Zn	107.06		
10	Zn	94.82		
13	Zn	100.75	138.38	
13	Zn	202.64		
13	Zn	111.76		
8	Zr	259.25	261.89	244.00
8	Zr	398.02		
8	Zr	128.40		
10	Zr	358.22	199.42	
10	Zr	173.39		
10	Zr	66.64		
13	Zr		270.69	
13	Zr	107.82		
13	Zr	433.56		

18 week weathering cycle: totals (aqua regia/HF)			TRT5	
Element	Container	Concentration	Triplicate average	Replicate average
			mg kg ⁻¹	
1	Al	89062.67	90177.83	89041.84
1	Al	83124.44		
1	Al	98346.39		
11	Al	92557.78	91010.48	
11	Al	85198.76		
11	Al	95274.89		
12	Al	82964.02	85937.21	
12	Al	91170.27		
12	Al	83677.34		
1	Ba	621.01	577.45	575.81
1	Ba	533.89		
11	Ba			
11	Ba			
11	Ba			
12	Ba	573.80	574.18	
12	Ba	584.29		
12	Ba	564.45		
1	Ca	4303.45	4306.66	4229.28
1	Ca	4309.87		
11	Ca			
11	Ca			
11	Ca			
12	Ca	4069.52	4151.90	
12	Ca	4320.60		
12	Ca	4085.57		
1	Cd	2.00	2.00	2.00
1	Cd	2.00		
1	Cd	2.00		
11	Cd	2.00	2.00	
11	Cd	2.00		
11	Cd	2.00		
12	Cd	2.00	2.00	
12	Cd	2.00		
12	Cd	2.00		
1	Co	27.46	27.28	25.26
1	Co	20.91		
1	Co	33.49		
11	Co	29.81	28.35	
11	Co	26.17		
11	Co	29.08		
12	Co	20.76	20.15	
12	Co	21.37		
12	Co	18.32		
1	Cr	80.93	80.04	78.49
1	Cr	75.59		
1	Cr	83.61		
11	Cr	77.88	76.93	
11	Cr	71.58		
11	Cr	81.32		
12	Cr	75.78	78.52	
12	Cr	83.23		
12	Cr	76.54		
1	Cu	82.19	79.59	65.19
1	Cu	74.45		
1	Cu	82.14		
11	Cu	55.67	54.82	
11	Cu	52.09		
11	Cu	56.70		
12	Cu	59.10	61.15	
12	Cu	66.06		
12	Cu	58.31		
1	Fe	41375.43	41143.45	40997.73
1	Fe	39134.90		
1	Fe	42920.04		
11	Fe	40656.65	40631.27	
11	Fe	38871.40		
11	Fe	42365.76		
12	Fe	40160.05	41218.47	
12	Fe	43049.45		

12	Fe	40445.90		
1	K	6418.49	5519.90	8073.03
1	K	4621.31		
1	K			
11	K	12130.94	12772.78	
11	K	12002.57		
11	K	14184.85		
12	K	6097.56	5926.40	
12	K	6996.15		
12	K	4685.49		
1	Mg	5823.39	5603.96	5551.11
1	Mg	5384.52		
1	Mg			
11	Mg			
11	Mg			
11	Mg			
12	Mg	5497.71	5498.26	
12	Mg	5876.92		
12	Mg	5120.14		
1	Mn	182.56	175.25	164.40
1	Mn	167.94		
1	Mn			
11	Mn			
11	Mn			
11	Mn			
12	Mn	148.96	153.55	
12	Mn	162.32		
12	Mn	149.38		
1	Mo	5.00	5.00	5.00
1	Mo	5.00		
1	Mo	5.00		
11	Mo	5.00	5.00	
11	Mo	5.00		
11	Mo	5.00		
12	Mo	5.00	5.00	
12	Mo	5.00		
12	Mo	5.00		
1	Na	7150.05	11131.04	12535.04
1	Na	15112.04		
1	Na			
11	Na			
11	Na			
11	Na			
12	Na	6334.94	13939.04	
12	Na	18170.64		
12	Na	17311.53		
1	Ni	35.57	29.68	29.37
1	Ni	26.16		
1	Ni	27.31		
11	Ni	21.57	28.30	
11	Ni	29.15		
11	Ni	34.20		
12	Ni	29.84	30.14	
12	Ni	39.47		
12	Ni	21.11		
1	P	1259.11	1289.43	1251.88
1	P	1240.49		
1	P	1368.70		
11	P	1232.16	1213.70	
11	P	1145.35		
11	P	1263.59		
12	P	1222.14	1252.50	
12	P	1282.44		
12	P	1252.94		
1	Pb	170.67	183.70	179.78
1	Pb	151.06		
1	Pb	229.38		
11	Pb	178.41	205.62	
11	Pb	158.19		
11	Pb	280.25		
12	Pb	162.55	150.02	
12	Pb	150.73		
12	Pb	136.78		
1	S	2328.90	2695.91	2699.37

1	S	2183.64		
1	S	3575.19		
11	S	3045.22	3048.20	
11	S	2972.29		
11	S	3127.09		
12	S	2322.00	2354.00	
12	S	2432.73		
12	S	2307.27		
1	Si	237066.67	312355.56	320704.59
1	Si	350000.00		
1	Si	350000.00		
11	Si	350000.00	337679.28	
11	Si	313037.84		
11	Si	350000.00		
12	Si	236236.79	312078.93	
12	Si	350000.00		
12	Si	350000.00		
1	Sr	252.58	333.46	323.46
1	Sr	228.88		
1	Sr	518.93		
11	Sr	383.89	399.51	
11	Sr	326.85		
11	Sr	487.79		
12	Sr	230.53	237.42	
12	Sr	254.78		
12	Sr	226.95		
1	Ti	2733.42	2745.85	2863.90
1	Ti	2515.32		
1	Ti	2988.80		
11	Ti	3064.70	2982.44	
11	Ti	2805.65		
11	Ti	3076.97		
12	Ti	2775.36	2863.41	
12	Ti	2995.97		
12	Ti	2818.91		
1	Zn	120.14	123.39	131.77
1	Zn	107.98		
1	Zn	142.07		
11	Zn	205.09	166.13	
11	Zn	170.82		
11	Zn	122.48		
12	Zn	102.74	105.78	
12	Zn	112.36		
12	Zn	102.25		
1	Zr	159.67	218.53	255.06
1	Zr	373.32		
1	Zr	122.61		
11	Zr	136.64	198.55	
11	Zr	103.55		
11	Zr	355.48		
12	Zr	170.80	348.10	
12	Zr	447.89		
12	Zr	425.63		

Appendix E.
Sequential-Selective Dissolution

2 week plot: SSD and residual									
elem	trt	KNO3	DI	NaOH	EDTA	HNO3	residual		Sum %
		% total (HF)	% total (HF)	% total (HF)	% total (HF)	% total (HF)	% total (HF)	% total (HF)	
Ba	0:0	0.199	0.052	0.013	11.818	1.286	86.633	100.000	
Ba	197:0	0.492	0.059	0.011	13.556	1.311	84.571	100.000	
Ba	197:197	3.378	0.007	0.008	13.929	3.015	79.663	100.000	
Ba	197:295	4.729	0.001	0.005	14.583	3.444	77.238	100.000	
Ba	197:394	5.724	0.033	0.003	16.132	4.011	74.096	100.000	
Co	0:0	13.662	1.981	1.207	6.592	10.644	65.914	100.000	
Co	197:0	11.637	1.193	1.248	7.561	9.306	69.055	100.000	
Co	197:197	6.663	0.956	1.923	6.108	9.140	75.209	100.000	
Co	197:295	5.641	0.963	1.789	6.799	9.666	75.142	100.000	
Co	197:394	4.429	0.898	2.350	6.485	9.590	76.247	100.000	
Cr	0:0	0.356	0.531	0.392	3.807	4.512	90.402	100.000	
Cr	197:0	0.383	0.551	0.389	1.177	4.809	92.690	100.000	
Cr	197:197	0.347	0.501	0.337	0.845	4.617	93.352	100.000	
Cr	197:295	0.338	0.499	0.329	1.155	4.908	92.772	100.000	
Cr	197:394	0.346	0.500	0.337	1.145	5.172	92.501	100.000	
Cu	0:0	68.739	0.198	11.096	14.086	5.880	0.000	100.000	
Cu	197:0	18.432	0.136	36.559	37.294	7.578	0.000	100.000	
Cu	197:197	12.807	0.030	24.261	25.416	8.564	28.921	100.000	
Cu	197:295	7.694	0.008	22.412	24.951	9.130	35.805	100.000	
Cu	197:394	5.936	0.091	20.817	23.063	10.141	39.953	100.000	
Mn	0:0	49.928	0.000	0.087	6.697	9.241	34.047	100.000	
Mn	197:0	37.355	0.000	0.093	7.283	7.005	48.285	100.000	
Mn	197:197	47.523	0.000	0.153	9.410	9.959	32.955	100.000	
Mn	197:295	36.773	0.000	0.160	8.785	8.549	45.734	100.000	
Mn	197:394	27.919	0.000	0.086	7.923	7.425	56.647	100.000	
Ni	0:0	21.823	3.527	3.161	17.939	20.425	33.125	100.000	
Ni	197:0	33.565	4.763	4.980	31.168	25.523	0.000	100.000	
Ni	197:197	11.596	2.485	3.710	17.209	14.852	50.149	100.000	
Ni	197:295	8.567	2.214	2.603	16.534	14.017	56.065	100.000	
Ni	197:394	7.324	2.162	3.245	16.901	14.672	55.696	100.000	
Pb	0:0	2.663	1.398	0.884	20.079	2.723	72.254	100.000	
Pb	197:0	1.018	1.447	1.000	20.868	2.501	73.166	100.000	
Pb	197:197	0.951	1.347	1.443	18.908	3.563	73.788	100.000	
Pb	197:295	0.913	1.298	1.127	18.918	3.839	73.904	100.000	
Pb	197:394	0.856	1.218	1.625	19.804	4.249	72.248	100.000	
Sr	0:0	0.079	0.146	0.039	25.878	0.369	73.489	100.000	
Sr	197:0	1.895	0.012	0.038	21.726	0.257	76.072	100.000	
Sr	197:197	4.401	0.000	0.024	20.167	4.839	70.569	100.000	
Sr	197:295	4.818	0.000	0.019	19.064	5.424	70.674	100.000	
Sr	197:394	5.108	0.000	0.018	18.379	6.062	70.402	100.000	
Zn	0:0	17.167	0.089	0.383	10.119	14.388	57.854	100.000	
Zn	197:0	16.779	0.000	0.414	12.037	14.191	56.580	100.000	
Zn	197:197	14.995	0.015	0.477	14.445	14.008	56.061	100.000	
Zn	197:295	12.728	0.000	0.482	15.248	12.923	58.619	100.000	
Zn	197:394	11.134	0.282	0.566	16.091	13.149	58.778	100.000	

Zn	4	10	11.073	0.000	0.438	15.204	11.580	61.705	100.000
Zn	4	10							0.000
Zn	4	13	13.144	0.000	0.929	17.937	13.948	54.042	100.000
Zn	4	13	11.159	0.000	0.369	14.863	10.800	63.009	100.000
Zn	4	13	15.010	0.000	0.806	16.092	13.186	54.905	100.000
Zn	5	1	12.779	0.000	0.544	15.457	12.949	58.270	100.000
Zn	5	1	10.870	0.000	0.713	14.769	11.558	62.091	100.000
Zn	5	1	12.499	0.000	0.500	17.612	16.367	53.022	100.000
Zn	5	11	10.043	0.000	0.676	15.781	18.646	54.854	100.000
Zn	5	11	11.242	0.000	0.721	13.439	12.045	62.553	100.000
Zn	5	11	11.015	2.512	0.397	16.927	11.165	57.984	100.000
Zn	5	12	9.037	0.000	0.371	14.620	10.270	65.702	100.000
Zn	5	12	10.577	0.074	0.614	18.293	12.295	58.147	100.000
Zn	5	12	12.151	0.000	0.530	17.658	12.793	56.867	100.000

Zn	4	10	10.604	0.000	0.420	14.559	11.089	36.672	95.760	59.088	61.705
Zn	4	10	10.901	0.000	0.407	15.706	11.151	38.165			
Zn	4	13	11.616	0.000	0.821	15.852	12.327	40.616	88.375	47.759	54.042
Zn	4	13	11.679	0.000	0.387	15.345	11.302	38.713	104.655	65.942	63.009
Zn	4	13	12.289	0.000	0.660	13.175	10.795	36.919	81.870	44.951	54.905
Zn	5	1	11.764	0.000	0.501	14.230	11.921	38.416	92.060	53.644	58.270
Zn	5	1	12.095	0.000	0.793	16.434	12.861	42.183	111.275	69.092	62.091
Zn	5	1	11.323	0.000	0.453	15.955	14.827	42.557	90.590	48.033	53.022
Zn	5	11	10.431	0.000	0.702	16.390	19.366	46.888	103.860	56.972	54.854
Zn	5	11	10.959	0.000	0.702	13.100	11.741	36.502	97.475	60.973	62.553
Zn	5	11	10.689	2.438	0.385	16.427	10.835	40.775	97.045	56.270	57.984
Zn	5	12	10.298	0.000	0.423	16.659	11.703	39.083	113.950	74.867	65.702
Zn	5	12	9.868	0.069	0.573	17.068	11.471	39.049	93.300	54.251	58.147
Zn	5	12	11.492	0.000	0.502	16.699	12.099	40.791	94.570	53.779	56.867

18 week weathering cycle: sequential selective dissolution								
		KNO3	Di water	NaOH	EDTA	HNO3	Residual	
elem	trf	% total	% total	% total	% total	% total	% total	Sum%
Ba	0:0	0.255	0.038	0.014	13.822	1.065	84.807	100.000
Ba	197:0	0.385	0.272	0.000	15.907	1.138	82.297	100.000
Ba	197:197	0.827	0.172	0.002	18.296	2.164	78.540	100.000
Ba	197:295	0.943	0.275	0.005	17.461	2.447	78.869	100.000
Ba	197:394	1.136	0.184	0.003	17.550	2.802	78.325	100.000
Co	0:0	2.768	1.301	0.955	8.858	11.099	75.018	100.000
Co	197:0	1.124	1.613	1.752	17.398	14.321	63.793	100.000
Co	197:197	0.870	1.248	1.643	13.731	12.785	69.723	100.000
Co	197:295	0.768	1.110	1.398	12.036	11.273	73.415	100.000
Co	197:394	0.594	0.855	1.388	10.991	8.694	77.478	100.000
Cr	0:0	0.365	0.524	0.354	6.338	5.170	87.249	100.000
Cr	197:0	0.370	1.182	0.312	1.080	6.201	90.855	100.000
Cr	197:197	0.366	1.109	0.311	1.408	6.467	90.338	100.000
Cr	197:295	0.439	1.152	0.288	1.213	5.772	91.137	100.000
Cr	197:394	0.318	0.964	0.270	1.136	5.334	91.977	100.000
Cu	0:0	21.041	0.000	13.570	21.434	10.864	33.092	100.000
Cu	197:0	0.235	0.444	13.039	25.544	5.547	55.192	100.000
Cu	197:197	0.215	0.550	13.449	26.836	6.994	51.956	100.000
Cu	197:295	0.195	0.514	12.601	24.827	7.151	54.711	100.000
Cu	197:394	0.153	0.403	10.658	20.689	6.092	62.004	100.000
Mn	0:0	4.732	0.000	0.038	6.351	4.192	84.687	100.000
Mn	197:0	0.403	0.230	0.061	56.125	9.427	33.754	100.000
Mn	197:197	0.583	0.300	0.055	69.366	11.798	17.898	100.000
Mn	197:295	0.768	0.257	0.038	65.047	11.459	22.431	100.000
Mn	197:394	0.459	0.187	0.068	50.373	8.691	40.221	100.000
Ni	0:0	6.554	6.271	4.321	28.977	35.792	18.085	100.000
Ni	197:0	3.496	5.015	7.904	53.792	29.793	0.000	100.000
Ni	197:197	2.820	4.089	6.382	45.177	26.461	15.070	100.000
Ni	197:295	2.168	3.133	4.834	36.408	19.748	33.710	100.000
Ni	197:394	1.701	2.456	3.901	30.645	16.100	45.197	100.000
Pb	0:0	2.029	1.080	0.782	19.101	3.078	73.929	100.000
Pb	197:0	0.978	1.386	1.177	21.888	3.656	70.914	100.000
Pb	197:197	0.828	1.188	1.141	18.641	3.728	74.474	100.000
Pb	197:295	0.822	1.188	1.026	18.535	3.805	74.624	100.000
Pb	197:394	0.695	1.001	1.035	16.336	3.221	77.711	100.000
Sr	0:0	0.081	0.034	0.026	16.232	0.376	83.251	100.000
Sr	197:0	0.952	0.004	0.032	17.011	0.500	81.500	100.000
Sr	197:197	2.868	0.004	0.022	19.141	3.076	74.889	100.000
Sr	197:295	3.119	0.000	0.026	16.576	3.642	76.638	100.000
Sr	197:394	2.690	0.000	0.012	13.061	3.184	81.053	100.000
Zn	0:0	5.721	0.027	0.409	10.493	17.600	65.751	100.000
Zn	197:0	0.195	0.690	0.363	28.507	17.815	52.431	100.000
Zn	197:197	0.194	0.674	0.574	31.600	18.083	48.875	100.000
Zn	197:295	0.270	0.593	0.652	27.824	14.890	55.771	100.000
Zn	197:394	0.155	0.492	0.628	26.770	12.715	59.239	100.000

Zn	4	10	0.187	0.559	0.527	32.388	16.302	50.037	100.000
Zn	4	10	1.292	0.704	0.685	34.333	18.008	44.978	100.000
Zn	4	13	0.199	0.673	0.575	34.003	17.545	47.006	100.000
Zn	4	13	0.099	0.305	0.268	17.500	6.685	75.143	100.000
Zn	4	13	0.179	0.548	0.365	30.953	14.127	53.828	100.000
Zn	5	1	0.166	0.578	0.456	30.717	12.963	55.119	100.000
Zn	5	1	0.185	0.584	0.660	32.973	14.813	50.785	100.000
Zn	5	1	0.141	0.477	0.418	27.159	11.441	60.364	100.000
Zn	5	11	0.098	0.376	0.336	16.037	8.169	74.984	100.000
Zn	5	11	0.117	0.339	0.493	18.656	10.551	69.844	100.000
Zn	5	11	0.163	0.367	0.611	33.002	13.445	52.412	100.000
Zn	5	12	0.195	0.639	1.541	30.648	17.101	49.876	100.000
Zn	5	12	0.178	0.762	0.681	29.999	14.565	53.814	100.000
Zn	5	12	0.235	0.508	0.947	35.218	17.399	45.693	100.000

Zn	4	10	0.200	0.599	0.564	34.675	17.453	53.491	107.060	53.569	50.037
Zn	4	10	1.225	0.667	0.650	32.553	17.074	52.169	94.815	42.646	44.978
Zn	4	13	0.200	0.678	0.579	34.256	17.676	53.389	100.745	47.356	47.006
Zn	4	13	0.200	0.618	0.543	35.461	13.546	50.368	202.635	152.267	75.143
Zn	4	13	0.200	0.613	0.408	34.593	15.789	51.602	111.760	60.158	53.828
Zn	5	1	0.200	0.694	0.548	36.904	15.574	53.920	120.140	66.220	55.119
Zn	5	1	0.200	0.630	0.712	35.603	15.995	53.140	107.975	54.835	50.785
Zn	5	1	0.200	0.677	0.594	38.584	16.254	56.309	142.065	85.756	60.364
Zn	5	11	0.200	0.772	0.689	32.889	16.753	51.304	205.085	153.781	74.984
Zn	5	11	0.200	0.579	0.843	31.867	18.024	51.513	170.820	119.307	69.844
Zn	5	11	0.200	0.450	0.748	40.421	16.467	58.286	122.480	64.194	52.412
Zn	5	12	0.200	0.656	1.583	31.488	17.589	51.497	102.740	51.243	49.876
Zn	5	12	0.200	0.856	0.765	33.708	16.365	51.892	112.955	60.463	53.814
Zn	5	12	0.240	0.520	0.969	36.009	17.789	55.526	102.245	46.719	45.693

Vita

Darlene Elizabeth Allred was born on May 22, 1972 in West Palm Beach, Florida. She graduated as valedictorian from Oliver Springs High School in 1990. She accepted an academic scholarship to Tennessee Technological University and studied in the College of Engineering until 1992. In January of 1993, she transferred to The University of Tennessee, Knoxville to study soils and the physical sciences. While studying towards her Bachelor of Science degree, she acquired membership to Phi Kappa Phi National Honor Society. In December of 1994, she graduated Magna Cum Laude with a Bachelor of Science Degree in Agriculture. Spring term of 1995, she accepted a graduate research assistantship at The University of Tennessee, Knoxville and began study towards the Master of Science degree in Plant and Soil Science, Soil Chemistry. Employment plans were undecided during the preparation of this vita.

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