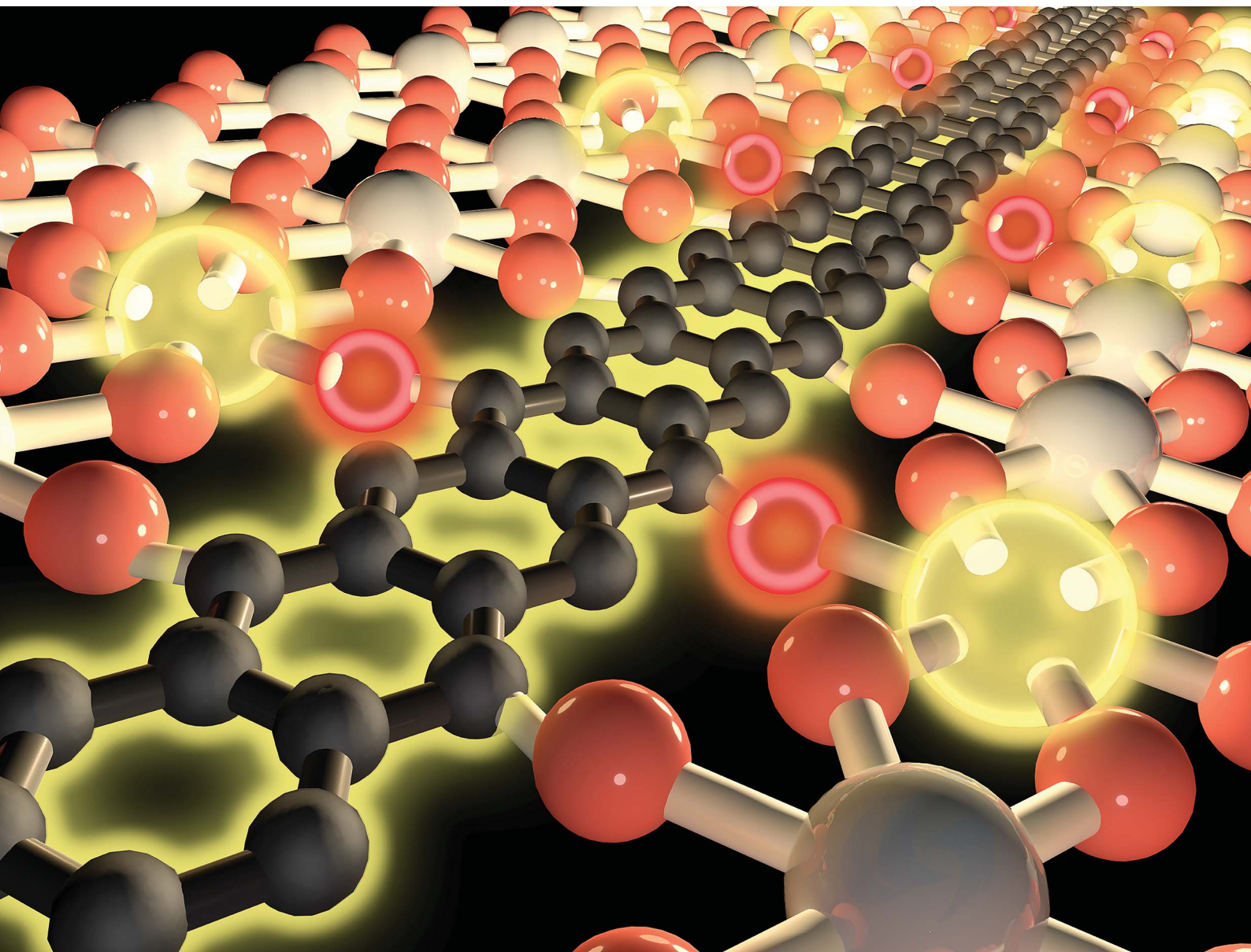


# Chemical Science

Volume 13  
Number 24  
28 June 2022  
Pages 7069–7366

rsc.li/chemical-science



ISSN 2041-6539

**EDGE ARTICLE**

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Cite this: *Chem. Sci.*, 2022, 13, 7126

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# Design and synthesis of TiO<sub>2</sub>/C nanosheets with a directional cascade carrier transfer†

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Directed transfer of carriers, akin to excited charges in photosynthesis, in semiconductors by structural design is challenging. Here, TiO<sub>2</sub> nanosheets with interlayered sp<sup>2</sup> carbon and titanium vacancies are obtained by low-temperature controlled oxidation calcination. The directed transfer of carriers from the excited position to Ti-vacancies to interlayered carbon is investigated and proven to greatly increase the charge transport efficiency. The TiO<sub>2</sub>/C obtained demonstrates excellent photocatalytic and photoelectrochemical activity and significant lithium/sodium ion storage performance. Further theoretical calculations reveal that the directional excited position/Ti-vacancies/interlayered carbon facilitate the spatial inside-out cascade electron transfer, resulting in high charge transfer kinetics.

Received 31st March 2022  
Accepted 9th May 2022

DOI: 10.1039/d2sc01872a

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## 1. Introduction

Semiconductors are able to generate excited carriers that diffuse to the surface for photo/electro chemical reactions. The fast transfer and efficient utilization of carriers are critical to high performance in the desired application.<sup>1–4</sup> Generally, the electron transfer at the single nanocrystal or interface of nanocrystals is random, often resulting in the electron dissipation and/or recombination before the electrons reach the surface of semiconductors for the redox reaction.<sup>5,6</sup> Therefore, directed transfer of carriers, akin to excited charges in photosynthesis, enables a minimum of energy-loss and a maximum of energy-

utilization. However, realizing directional charge transfer in semiconductors by structural design is challenging.

Titanium oxide (TiO<sub>2</sub>) has gained great attention as one of the most promising semiconductors for various photo and electrochemical applications.<sup>7–8</sup> Carbon has been widely used to composite TiO<sub>2</sub> as an electron acceptor for enhancement of the charge transfer.<sup>9–11</sup> Carbonization in an inert atmosphere is a common way to fabricate the composition and doping structure.<sup>12,13</sup> However, it will inevitably also form undesired carbon species such as low conductive sp<sup>3</sup> carbon on the surface. Nanocarbon with a rich sp<sup>2</sup> carbon structure, such as graphene sheets, enhances the conductivity and enables efficient charge transfer conductive performance in TiO<sub>2</sub>-based composites.<sup>14,15</sup> However, the directed transfer of electrons only occurs in a limited range of the interfacial junction between TiO<sub>2</sub> and nanocarbon in most cases. Recently, we have designed a spatially inside-out heterojunction with directed transfer of carriers, in which CQDs act as a bridge for interfacial charge transfer and graphene as a net for charge collection.<sup>11</sup> It has to be pointed out that this ternary structure requires very careful adjustment and pre-formed nanocarbon with the sp<sup>2</sup> structure. At high calcination temperature (over 1000 °C),<sup>16</sup> the hydrocarbon precursor starts to transform into the graphitic sp<sup>2</sup> structure, but at this temperature the size of the semiconductor particles, their crystal phases and morphologies would be greatly changed and even fused. Therefore, the common calcination method in an inert atmosphere is not suitable for the preparation of rich sp<sup>2</sup> carbon in TiO<sub>2</sub> via precursor carbonization. Note that the calcination in air (so-called oxidative calcination) could easily lead to the removal of the surface carbon species of nanocomposites by oxidation at mild temperature (300–600 °C),<sup>17</sup> but not to the removal of the carbon

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† Electronic supplementary information (ESI) available. See <https://doi.org/10.1039/d2sc01872a>





species from the crystal framework. Therefore, one theoretical possibility arises, namely that the oxidative calcination not only transforms titanium compounds into  $\text{TiO}_2$  and remove the unsatisfied surface carbon, but also forms inner rich- $\text{sp}^2$  carbon confined by  $\text{TiO}_2$ . The nanosheet structure recently developed is very elegant, which enables an abundance of vacancies and coordinatively unsaturated sites, shortens the mass diffusion length and greatly improves charge transfer performance.<sup>18–20</sup> The rich- $\text{sp}^2$  carbon structure in  $\text{TiO}_2$  nanosheets would therefore be a very effective design for directional charge transfer.

In this study, lamellar titanium glycerolate is chosen as the precursor for the synthesis of anatase  $\text{TiO}_2$  nanosheets with interlayered rich- $\text{sp}^2$  carbon using a well-defined oxidization calcination process. The oxygen-rich environment contributes to the formation of titanium vacancies (Ti-vacancies). The directed transfer of carriers from the excited position to Ti-vacancies to interlayered carbon is investigated and proven to greatly increase the charge transport efficiency and electrical conductivity. Consequently,  $\text{TiO}_2/\text{C}$  demonstrates significant photocatalytic and photoelectrochemical activity and excellent lithium/sodium ion storage.

## 2. Results and discussion

TGA-DTA of titanium glycerolate (denoted as Ti-G) (Fig. S1, ESI†) indicates the removal of surface carbon in the range of 300–339 °C and the phase transformation from anatase to the rutile structure at 435 °C. Therefore, we choose 350 °C for the synthesis of the metastable anatase phase  $\text{TiO}_2/\text{C}$  composite without surface carbon, while the interlayered carbon species are still coordinated to  $\text{TiO}_2$  (detailed description in the ESI†). SEM images (Fig. S2a, ESI†) show that Ti-G is a flower-like particle with a uniform size of around 3  $\mu\text{m}$ , and calcination leads to aggregation of nanocrystals in the branches (Fig. S2, ESI†). XRD patterns (Fig. S3, ESI†) and Raman spectra (Fig. S4, ESI†) indicate that calcination at 350 °C in air results in the phase transformation from titanium glycerolate to anatase  $\text{TiO}_2$ . And calcination at 550 °C could lead to the formation of the rutile phase.  $\text{N}_2$  adsorption-desorption measurements (Fig. S5, ESI†) of the  $\text{TiO}_2$  sample calcined at 350 °C in air disclose a type II isotherm (indicative of non- or macroporous adsorbents) with an H3 hysteresis loop (due to non-rigid aggregates of plate-like particles).<sup>21</sup> It also shows mesopores of around 3 nm and a large fraction of mesopores larger than 10 nm by the packing of the  $\text{TiO}_2$  nanocrystals, with a tube structure of 25 nm. Notably, the sample calcined at 350 °C in Ar has only 24  $\text{m}^2 \text{g}^{-1}$  of BET surface area with no obvious porous structure (Fig. S5 and Table S1, ESI†), because it has surface carbon that could block the pores.

As shown in Fig. 1a and the inset, the flower-like structure of  $\text{TiO}_2/\text{C}_{\text{inter}}$  is assembled by very thin nanosheets. The STEM image and the corresponding EDX spectral mapping images (Fig. 1b–e) indicate the uniform distribution of Ti, C and O elements. The AFM tomography studies (Fig. 1f) show that the branches of the particle are nanosheets with a thickness of around 4–5 nm, which is in accordance with the magnified TEM image (Fig. S6, ESI†). The HRTEM image (Fig. 1h, S7a and b,

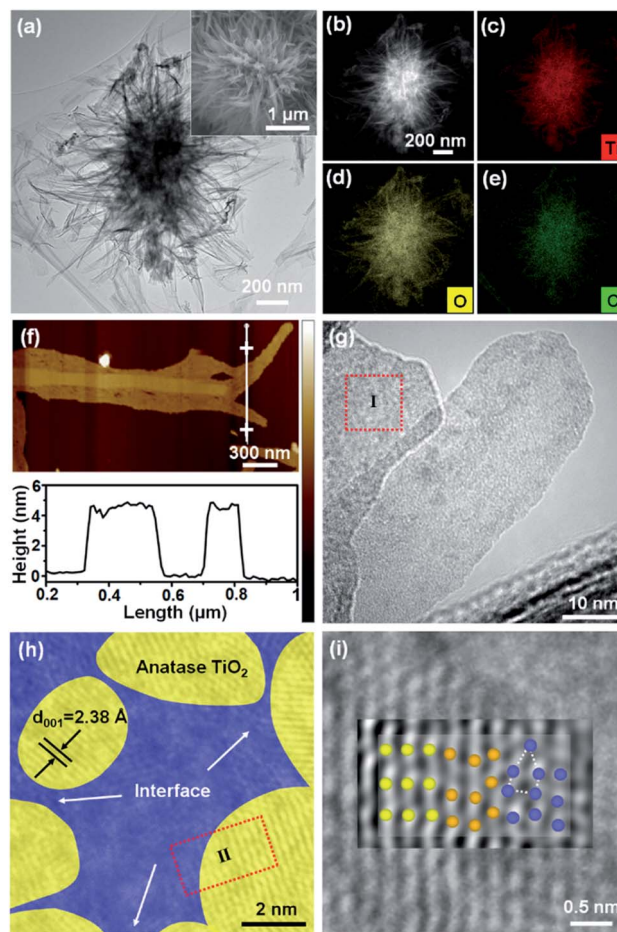


Fig. 1 (a) TEM image, inset: SEM image, (b) STEM image and the corresponding EDX mapping image of (c) Ti, (d) O, and (e) C elements, (f) AFM topography image and the corresponding height information of  $\text{TiO}_2/\text{C}_{\text{inter}}$  (g) TEM image, (h) magnified TEM image of region I and (i) original image and inverse FFT image of region II, and the corresponding atomic models of the highly crystalline phase (ordered lattice, light yellow, left), nanofusion phase (dark yellow, middle), and semi-crystalline/amorphous phase (disordered defects, blue, right) of  $\text{TiO}_2/\text{C}_{\text{inter}}$ . Fig. parts (h) and (i) are given in Fig. S7, ESI† as originals without the interpreting color overlay.

ESI†) shows the lattice fringes with an interplanar spacing of 0.35 nm, which agrees with the (101) planes of anatase  $\text{TiO}_2$ . The anatase  $\text{TiO}_2$  nanocrystals (Fig. 1h, yellow areas) are surrounded by an amorphous phase (Fig. 1h, blue areas). In the HRTEM image and the inverse FFT image (Fig. 1i and S7b–d, ESI†), the ordered lattice fringes (light yellow, left), the lattice distortions (dark yellow, middle) and the disordered amorphous phase (blue, right, probably involving amorphous  $\text{TiO}_2$  and carbon) are clearly recognized. This kind of crystalline/semi-crystalline/amorphous interface is coherent at the atomic scale and thus becomes a platform for the inner carbon and defect generation and minimizes interface loss of energy.

<sup>13</sup>C cross-polarization/magic-angle spinning NMR is an effective method to investigate carbon species comprehensively and quantitatively.<sup>22–24</sup> For comparison,  $\text{TiO}_2$  mainly with surface carbon (named  $\text{TiO}_2/\text{C}_{\text{surf}}$ ) by calcination from Ti-G at



350 °C in an Ar atmosphere) is also characterized. Ti-G shows three well resolved peaks between 69 and 90 ppm (Fig. S8, ESI<sup>†</sup>), indicating the successful coordination of glycerol.<sup>25,26</sup> The peaks near 138 ppm in TiO<sub>2</sub>/C<sub>Inter</sub> (Fig. 2a and S9, ESI<sup>†</sup>) and 135 ppm in TiO<sub>2</sub>/C<sub>Surf</sub> are assigned to sp<sup>2</sup> hybridized C atoms.<sup>24,27</sup> The peaks around 30–50 ppm are assigned to sp<sup>3</sup> C atoms.<sup>27</sup> Besides, TiO<sub>2</sub>/C<sub>Inter</sub> shows an additional peak at around 180 ppm, which can be ascribed to C=O with sp<sup>2</sup> hybridized carbon. Very interestingly, the ratio of sp<sup>2</sup> and sp<sup>3</sup> in TiO<sub>2</sub>/C<sub>Inter</sub> is about 2.6-fold that of the TiO<sub>2</sub>/C<sub>Surf</sub> (Fig. S9 and Table S2, detailed description in the ESI<sup>†</sup>), suggesting a high level of sp<sup>2</sup> carbon in TiO<sub>2</sub>/C<sub>Inter</sub>. Raman spectra (Fig. S10, ESI<sup>†</sup>) show only graphite carbon in TiO<sub>2</sub>/C<sub>Inter</sub> and the weak peak indicates the low content and inner structure of carbon species.<sup>28</sup> The high-resolution C 1s spectra (Fig. 2b) could be deconvoluted as the C–C bond (around 284.8 eV), C–O–(Ti) bond (around 286.3 eV) and O–C=O (around 288.7 eV).<sup>29,30</sup> There is an obvious difference of the peak area ratio of the C–C bond and C–O–(Ti) bond between TiO<sub>2</sub>/C<sub>Inter</sub> (0.84) and TiO<sub>2</sub>/C<sub>Surf</sub> (4.52) (Fig. 2b). Further evidence is provided by the surface carbon content and C/Ti ratio calculated from Ti 2p and C 1s XPS data (Fig. S11 and Table S3, detailed description in the ESI<sup>†</sup>). TiO<sub>2</sub>/C<sub>Inter</sub> shows a low surface carbon content (15.0 at%) near to the sample calcined at 550 °C (15.5 at%), while TiO<sub>2</sub>/C<sub>Surf</sub> shows a relatively high carbon content (44.7 at%) similar to the original carbon content in Ti-G (47.0 at%). These observations suggest that the surface carbon content can be reduced by oxidization calcination but the inner carbon species in TiO<sub>2</sub>/C<sub>Inter</sub> would remain.

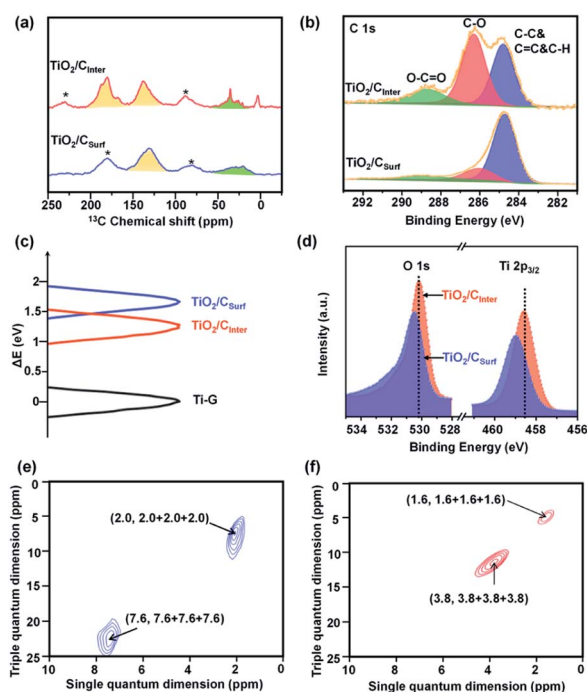


Fig. 2 (a) <sup>13</sup>C NMR spectra of TiO<sub>2</sub>/C<sub>Inter</sub> and TiO<sub>2</sub>/C<sub>Surf</sub>. \* indicates the signal of the rotor cap. (b) XPS C 1s spectra of TiO<sub>2</sub>/C<sub>Inter</sub> and TiO<sub>2</sub>/C<sub>Surf</sub>. (c) Ti 2p<sub>3/2</sub> chemical shift among Ti-G, TiO<sub>2</sub>/C<sub>Inter</sub> and TiO<sub>2</sub>/C<sub>Surf</sub>. (d) Ti 2p<sub>3/2</sub> and O 1s XPS spectra of TiO<sub>2</sub>/C<sub>Inter</sub> and TiO<sub>2</sub>/C<sub>Surf</sub>, and (e) and (f) <sup>1</sup>H TQ-SQ MAS NMR spectra of (e) TiO<sub>2</sub>/C<sub>Inter</sub>, and (f) TiO<sub>2</sub>/C<sub>Surf</sub>.

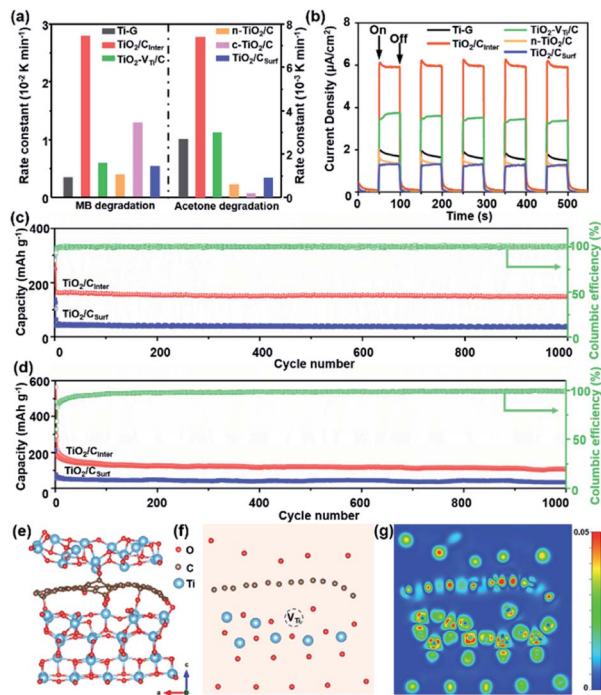
This can also be observed in the FT-IR spectra (Fig. S12, detailed description in the ESI<sup>†</sup>). After low temperature calcination, TiO<sub>2</sub>/C<sub>Inter</sub> exhibits a grey color with enhanced visible-light absorption and a narrower bandgap (Fig. S13, ESI<sup>†</sup>). The more positive Ti 2p chemical shifts of TiO<sub>2</sub>/C<sub>Inter</sub> and TiO<sub>2</sub>/C<sub>Surf</sub> indicate the increase of the electron–electron repulsion around the Ti atoms by strong interactions of the Ti–O–C bonds (Fig. 2c and d). The O 1s core-level XPS spectrum of TiO<sub>2</sub>/C<sub>Inter</sub> (Fig. S11c, ESI<sup>†</sup>) displays the major peaks caused by the O<sup>2–</sup> ions in the O–Ti–O lattice (around 530.0 eV).<sup>31,32</sup> The O<sup>–</sup> species (around 532.0 eV) are considered to correlate with the Ti-vacancies to compensate for the Ti<sup>4+</sup> deficiencies.<sup>33–35</sup> The formation of Ti-vacancies can be deduced from the surface O/Ti ratio of the samples (Fig. S11e and Table S3, detailed description in the ESI<sup>†</sup>). The EPR spectrum of TiO<sub>2</sub>/C<sub>Inter</sub> (Fig. S14, ESI<sup>†</sup>) shows a strong signal at *g* = 1.998, which could be assigned to Ti-vacancies.<sup>36,37</sup>

The two-dimensional (2D) <sup>1</sup>H TQ-SQ MAS NMR method is an effective way to identify the local structure and interactions in TiO<sub>2</sub> by the investigation of titanol (Ti–OH) sites.<sup>38,39</sup> The dipolar interactions of <sup>1</sup>H can directly probe the spatial proximities of Ti–OH groups such as surface Ti–OH, Ti–OH from broken Ti–O–Ti bonds and Ti–OH nests from Ti vacancies. More importantly, the interlayered carbon could also affect the chemical state of the Ti–OH groups. The signal of surface Ti–OH groups in mutual spatial proximity is observed in both TiO<sub>2</sub>/C<sub>Surf</sub> (2.0, 2.0 + 2.0 + 2.0) (Fig. 2e) and TiO<sub>2</sub>/C<sub>Inter</sub> (1.6, 1.6 + 1.6 + 1.6) (Fig. 2f). Besides, TiO<sub>2</sub>/C<sub>Surf</sub> shows a signal at (7.6, 7.6 + 7.6 + 7.6), which indicates that the Ti–OH species from broken Ti–O–Ti bonds are spatially close. Meanwhile, TiO<sub>2</sub>/C<sub>Inter</sub> shows a signal at (3.8, 3.8 + 3.8 + 3.8), which can be assigned to the Ti–OH nests caused by titanium vacancies. Notably, the chemical shift of the Ti–OH nests in TiO<sub>2</sub>/C<sub>Inter</sub> moves to a high field,<sup>35</sup> indicating that the interlayered carbon species have strong interactions with the Ti–OH nests. It can be deduced that low-temperature calcination in air could remove the surface carbon species and maintain the interlayered carbon, meanwhile contributing to the formation of titanium vacancies.

To further investigate the role of interlayered carbon in TiO<sub>2</sub>/C<sub>Inter</sub>, photodegradation of an organic pollutant in liquid and gas phases (methylene blue, MB and acetone, respectively) was performed. For comparison, the performance of Ti-G, TiO<sub>2</sub> with titanium vacancies and mixed carbon (named TiO<sub>2</sub>-V<sub>Ti</sub>/C), normal TiO<sub>2</sub> without titanium vacancies but with carbon (named n-TiO<sub>2</sub>/C), commercial TiO<sub>2</sub> nanotubes with carbon (named c-TiO<sub>2</sub>/C) and TiO<sub>2</sub>/C<sub>Surf</sub> was also investigated. TiO<sub>2</sub>/C<sub>Inter</sub> exhibits the highest photocatalytic activity (Fig. 3a) compared to other TiO<sub>2</sub>/C composites in MB and acetone degradation. The photocatalytic stability test of TiO<sub>2</sub>/C<sub>Inter</sub> shows that around 97% of photocatalytic activity is retained after five cycles of photocatalysis (Fig. S15a, ESI<sup>†</sup>), and a strong signal of Ti-vacancies from EPR can be clearly observed (Fig. S15b, ESI<sup>†</sup>), suggesting the high stability of Ti-vacancies. Photoelectrochemical studies (Fig. 3b) reveal that TiO<sub>2</sub>/C<sub>Inter</sub> exhibits the highest photocurrent intensity of 6 μA cm<sup>-2</sup>, which is 3-fold, 1.5-fold, 3.7-fold and 3.8-fold that of Ti-G, TiO<sub>2</sub>-V<sub>Ti</sub>/C, n-TiO<sub>2</sub>/C, and TiO<sub>2</sub>/C<sub>Surf</sub>, respectively. The EIS Nyquist plots of







**Fig. 3** (a) Photocatalytic rate constants for degradation of methylene blue (MB) and acetone with (a) Ti-G, (b)  $\text{TiO}_2/\text{C}_{\text{Inter}}$ , (c)  $\text{TiO}_2\text{-V}_{\text{Ti}}/\text{C}$ , (d) n- $\text{TiO}_2/\text{C}$  (e) c- $\text{TiO}_2/\text{C}$  and (f)  $\text{TiO}_2/\text{C}_{\text{Surf}}$ , and (b) transient photocurrent response of Ti-G,  $\text{TiO}_2/\text{C}_{\text{Inter}}$ ,  $\text{TiO}_2\text{-V}_{\text{Ti}}/\text{C}$ , n- $\text{TiO}_2/\text{C}$  and  $\text{TiO}_2/\text{C}_{\text{Surf}}$ . (c) Lithium-ion and (d) sodium-ion long cycle performance at a current density of 10 C. (e) The optimized model of  $\text{TiO}_2$  with interlayered carbon and titanium vacancies, (f) the section model of  $\text{TiO}_2/\text{C}_{\text{Inter}}$  of the (010) facet and (g) the corresponding charge density difference.

$\text{TiO}_2/\text{C}_{\text{Inter}}$  after irradiation with UV-Vis light display the smallest semicircle among all the samples (Fig. S16, ESI<sup>†</sup>), indicating that titanium vacancies and interlayered carbon are beneficial for the hole transfer to the electrolyte and thus greatly reduce the charge transfer resistance.<sup>40</sup>

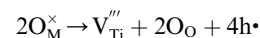
$\text{TiO}_2/\text{C}$  samples were further applied as anode materials for lithium/sodium storage properties. It can be observed that  $\text{TiO}_2/\text{C}_{\text{Inter}}$  exhibits the highest charge capacity after 300 cycles at 1 C in comparison with other  $\text{TiO}_2/\text{C}$  anodes (Fig. S17, ESI<sup>†</sup>) and good structural stability (Fig. S18, detailed description in the ESI<sup>†</sup>). Further,  $\text{TiO}_2/\text{C}_{\text{Inter}}$  has better reversible capacity and rate capability in comparison with  $\text{TiO}_2/\text{C}_{\text{Surf}}$  (Fig. S19, ESI<sup>†</sup>). At a current density of 10 C ( $1.7 \text{ A g}^{-1}$ ),  $\text{TiO}_2/\text{C}_{\text{Inter}}$  shows a high reversible specific capacity of  $148 \text{ mA h g}^{-1}$  for lithium storage and  $108 \text{ mA h g}^{-1}$  for sodium storage after 1000 cycles at 10 C (Fig. 3c and d), demonstrating its outstanding cycling performances in comparison with other  $\text{TiO}_2/\text{C}$ -based electrodes (Table S4, ESI<sup>†</sup>).

To gain further insight into the decisive role of interlayered carbon in efficient charge transfer and the stability of  $\text{TiO}_2/\text{C}_{\text{Inter}}$  from a theoretical point of view, density functional theory (DFT) calculations were performed. The charge density difference of an optimized model of  $\text{TiO}_2$  with interlayered carbon and titanium vacancies (Fig. 3e and S20a, ESI<sup>†</sup>) shows obvious

accumulation of electrons at the interfacial carbon layer in comparison with  $\text{TiO}_2$  with surface carbon species or no carbon (Fig. S20b–d, ESI<sup>†</sup>). The sectional charge density difference gives a clear view of the charge accumulation around the titanium vacancies and interlayer carbon, which could act as a bridge for efficient cascade and directed charge transfer from inside the lattice to the outer surface (Fig. 3f and g). The calculated formation energy (Tables S5 and S6<sup>†</sup>) indicates the good stability of  $\text{TiO}_2$  with interlayered carbon and titanium vacancies (detailed description of theoretical calculations in the ESI<sup>†</sup>). The calculation results clearly indicate that interlayered carbon can significantly facilitate the charge transfer and therefore enhance the photo/electro-catalytic performance.

The formation process of  $\text{TiO}_2/\text{C}_{\text{Inter}}$  we proposed is illustrated in Fig. 4. Ti-G with the oriented lamellar structure (Fig. 4a and b) starts to dehydrate and carbonize during the early stage of calcination at  $350^\circ\text{C}$ . The surface carbon and inner carbon form, which contain  $\text{sp}^2$  and  $\text{sp}^3$  types of carbon. With prolongation of the calcination time, the surface carbon is oxidized and removed (Fig. 4c). At the same time, the inner carbon remains and further rearranges to the  $\text{sp}^2$  type because of the confinement effect by the lamellar structure of  $\text{TiO}_2$  (Fig. 4d).

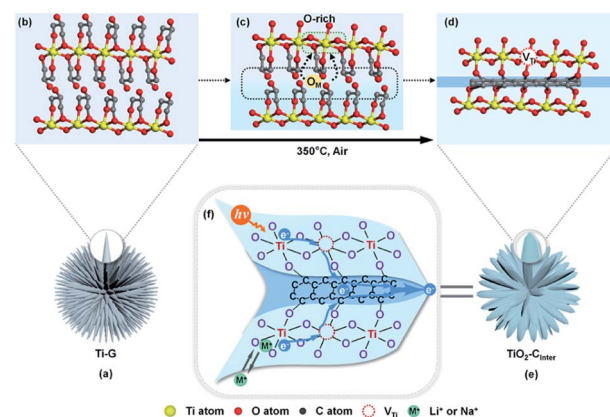
In the meantime, the escaped oxygen from the inner phase would form an oxygen-rich environment at the interface, which is beneficial for the formation of titanium vacancies (Fig. 4c and d). The formation of titanium vacancies can be described as:



where  $\text{V}_{\text{Ti}}^{\prime\prime}$  represents a titanium vacancy, and  $\text{h}^\bullet$  represents a hole.

Finally,  $\text{TiO}_2$  nanosheets with interlayered carbon and titanium vacancies (Fig. 4d and e) are obtained by this low temperature calcination strategy (detailed description of the formation of interfacial defects in the ESI<sup>†</sup>).

The proposed mechanism of  $\text{TiO}_2/\text{C}_{\text{Inter}}$  for photocatalysis and lithium/sodium storage is shown in Fig. 4f. The interlayer carbon and Ti-vacancies form a spatial inside-out electron



**Fig. 4** (a)–(e) Schematic illustration of the formation mechanism of  $\text{TiO}_2/\text{C}_{\text{Inter}}$ . (b)–(d) the simulated atomic change process from Ti-G to  $\text{TiO}_2/\text{C}_{\text{Inter}}$ ,  $\text{O}_M$  refers to migrating oxygen. (f) Schematic description of the fast electron-transfer pathway and the proposed mechanism.



transfer cascade from the lattice to the surface, which is not only beneficial to the charge separation in photocatalysis, but also enhances the interfacial conductivity for efficient electron transfer and  $\text{Li}^+/\text{Na}^+$  insertion. Besides, the ultrathin nanosheet could shorten the diffusion length of electrons (including photogenerated electrons) and  $\text{Li}^+/\text{Na}^+$ , and the high specific area of nanosheets provides large contact area for photo/electrocatalytic reactions. Moreover, the amorphous  $\text{TiO}_2$ /carbon interface is helpful to restrain the collapse of the nanostructure in the insertion/removal reactions, thus contributing to excellent stability.

### 3. Conclusions

In summary,  $\text{TiO}_2$  nanosheets with interfacial carbon and titanium vacancies have been prepared successfully by a controlled oxidation calcination. The confined carbon in the interlayer is crucial to a directional and efficient charge transfer, thus contributing to significantly increased photocatalytic activity and electrochemical performance. Our work provides a simple but effective way for high-performance design of semiconductors featuring low cost, high efficiency, and high stability.

### Data availability

The data that supports the findings of this study are available within the ESI† and from the corresponding author upon reasonable request.

### Author contributions

S. M. W. carried out the experiments of synthesis, photocatalytic performance and the lithium/sodium storage experiments. X. Y. Y. conceived the project, provided the idea, and guided the experiments. Y. T. W., Y. X. Z. S. T. X. helped with the experiments. X. F. Z. and L. Y. W. helped with the NMR measurements and corresponding analysis. S. M. W. and X. Y. Y. proposed the mechanisms. J. B. C. performed the DFT calculation and helped with the analysis. G. T. performed the measurements of TEM. S. M. W. and X. Y. Y. wrote and revised the paper. C. J., M. S. and D. W. B. revised the paper. All authors have given approval to the final version of the manuscript.

### Conflicts of interest

There are no conflicts to declare.

### Acknowledgements

This work was supported by a joint National Natural Science Foundation of China-Deutsche Forschungsgemeinschaft (NSFC-DFG) project (NSFC grant 51861135313, DFG JA466/39-1), Sino-German Center COVID-19 Related Bilateral Collaborative Project (C-0046), Shenzhen Science and Technology Program (JCYJ20210324142010029), Guangdong Basic and Applied Basic Research Foundation (2019A1515110435),

Guangdong Province International Scientific and Technological Cooperation Projects (2020A0505100036), National 111 project (B20002) and PCSIRT (IRT\_15R52). D.W.B. acknowledges financial support from Saint Petersburg State University (Research Grant 39054581). The authors would like to thank Prof. Reshef Tenne from Weizmann Institute of Science, Lu Wu from Hubei University for helpful discussion and the Nanostructure Research Centre (NRC) for the S/TEM work.

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