

ground emissions. In addition, signal transduction in the wavelength ratiometric mode is highly advantageous.

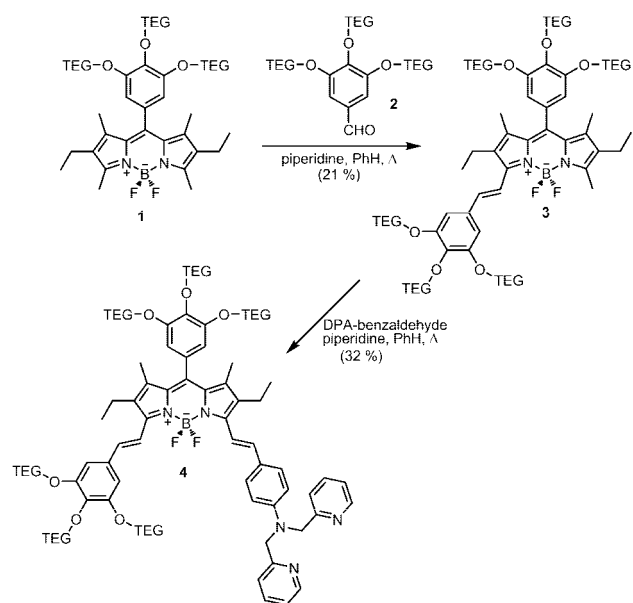
In recent literature, a large number of fluorescent probes for zinc have been proposed based on different fluorophores such as quinoline,⁵ boradiazaindacene (bodipy),⁶ dansyl,⁷ fluorescein,⁸ anthracene,⁹ acridine,¹⁰ benzofuran,¹¹ naphthalimide,¹² cyanine,¹³ and squaraines.¹⁴

Unfortunately, for most of them, water solubility is limited, thus various solvent mixtures are used in characterization. This fact, together with the requirement for shorter wavelength excitation, hinders biological applications.

As part of our ongoing interest in rational design of ratiometric fluorescent chemosensors, emitting especially in the red end of the visible spectrum, we are particularly intrigued by the distyryl–bodipy¹⁵ dyes. These novel derivatives of well-known bodipy dyes are emerging as promising red emitting dyes with interesting applications as sensitizers for photodynamic therapy^{15a} and ion sensing.^{15b} The distyryl–bodipy dyes can be synthesized with water solubilizing groups, which increases the likelihood of biological applications. When the styryl groups carry dialkylamino groups, the ICT process can be modulated by cation binding which provides ample opportunities of cation sensing.¹⁶

With these considerations, we set out to synthesize compound **4** (Scheme 1). The synthesis requires stepwise Knoevenagel condensations with two different aldehydes to yield the distyryl–bodipy. The TEG (triethyleneglycol) groups are placed on a gallic acid derived, 3,4,5-trihydroxyphenyl unit. The metal chelator is a well-known Zn(II) specific ligand, dipicolylamine (DPA). The target compound **4** was obtained in satisfactory yield and characterized by ¹H and ¹³C NMR and HRMS. Six TEG groups are clearly enough to make this molecule compatible with water, and the spectral

Scheme 1. Synthesis of the Near IR Emitting, Water-Soluble DS–Bodipy-Based Chemosensor **4**



data to be discussed from this point on are obtained in 5:95 ethanol–aqueous buffer solutions.

The absorption spectrum (Supporting Information) displays a broad peak centered at 680 nm, and the molar extinction coefficient is 72 000 M⁻¹cm⁻¹ at this wavelength. The emission spectrum obtained in aqueous buffer solution (5% ethanol was added as a cosolvent) shows a peak centered at 726 nm (Figure 1). Gradual addition of Zn(II) ions to this

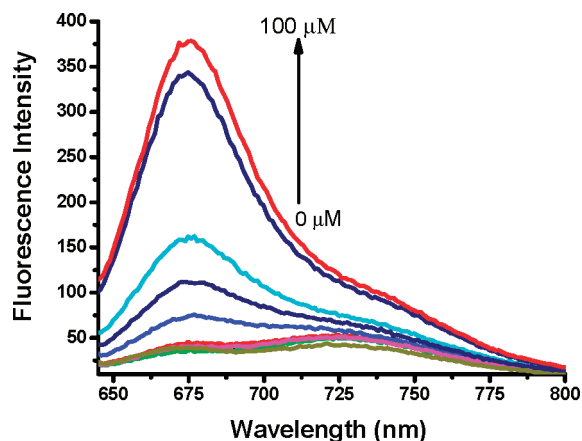


Figure 1. Change in the emission spectrum of the chemosensor **4** (2.5 μM) in response to increasing concentrations of Zn(II) in an ethanol–aqueous buffer mixture (5% in ethanol, HEPES 0.1 M, pH = 7.2). Zinc concentrations were varied in the following order: 0, 0.25, 0.5, 1, 2, 5, 10, 20, 50, 100 μM. Excitation was at 630 nm, with slit widths of 5 nm.

solution results in a blue shift to 625 nm with a concomitant increase in emission intensity. This clearly shows that the

(5) (a) Fahrni, C. J.; O'Halloran, T. V. *J. Am. Chem. Soc.* **1999**, *121*, 11448–11458. (b) Zhang, Y.; Guo, X.; Si, W.; Jia, L.; Qian, X. *Org. Lett.* **2008**, *10*, 473–476.

(6) (a) Turfan, B.; Akkaya, E. U. *Org. Lett.* **2002**, *4*, 2857–2859. (b) Harriman, A.; Mallon, L. J.; Stewart, B.; Ulrich, G.; Ziessel, R. *Eur. J. Org. Chem.* **2007**, 3191–3198.

(7) Jiang, P.; Chen, L.; Lin, J.; Liu, Q.; Ding, J.; Gao, X.; Guo, Z. *Chem. Commun.* **2002**, 1424–1425.

(8) Sun, W. C.; Gee, K. R.; Klaubert, D. H.; Haugland, R. P. *J. Org. Chem.* **1997**, *62*, 6469–6475.

(9) (a) Ojida, A.; Mito-Oka, Y.; Inoue, M.; Hamachi, I. *J. Am. Chem. Soc.* **2002**, *124*, 6256–6258. (b) Ertas, N.; Akkaya, E. U.; Ataman, O. Y. *Talanta* **2001**, *51*, 693–699.

(10) Park, M. S.; Swamy, K. M. K.; Lee, Y. J.; Lee, H. N.; Jang, Y. J.; Moon, Y. H.; Yoon, J. *Tetrahedron Lett.* **2006**, *47*, 8129–8132.

(11) Maruyama, S.; Kikuchi, K.; Hirano, T.; Urano, Y.; Nagano, T. *J. Am. Chem. Soc.* **2002**, *124*, 10650–10651.

(12) (a) Xu, Z.; Qian, X.; Cui, J.; Zhang, R. *Tetrahedron* **2006**, *62*, 10117. (b) Lee, T. C.; Parkesh, R. *Org. Biomol. Chem.* **2003**, *1*, 3265–3267. (c) Parkesh, R.; Lee, T. C.; Gunnlaugsson, T. *Org. Biomol. Chem.* **2007**, *5*, 310–317.

(13) Kiyose, K.; Kojima, H.; Urano, Y.; Nagano, T. *J. Am. Chem. Soc.* **2006**, *128*, 6548–6549.

(14) (a) Dilek, G.; Akkaya, E. U. *Tetrahedron Lett.* **2000**, *41*, 3721–3724. (b) Carol, P.; Sreejith, S.; Ajayaghosh, A. *Chem. Asian J.* **2007**, *2*, 338–348.

(15) (a) Dost, Z.; Atilgan, S.; Akkaya, E. U. *Tetrahedron* **2006**, *62*, 8484–8488. (b) Atilgan, S.; Ekmekci, Z.; Dogan, A. L.; Guç, D.; Akkaya, E. U. *Chem. Commun.* **2006**, 4398–4400.

(16) (a) Coskun, A.; Akkaya, E. U. *J. Am. Chem. Soc.* **2005**, *127*, 10464–10465. (b) Coskun, A.; Akkaya, E. U. *J. Am. Chem. Soc.* **2006**, *128*, 14474–14475. (c) Peng, X.; Du, J.; Fan, J.; Wang, J.; Wu, Y.; Zhao, J.; Sun, S.; Xu, T. *J. Am. Chem. Soc.* **2007**, *129*, 1500–1501.

coordination of Zn(II) ions effectively blocks excited-state charge transfer from the dipicolylphenyl group. Hill plot analysis of the data obtained in the titration of **4** with Zn(II) yielded a 1:1 stoichiometry (slope = 0.97) with a dissociation constant (K_d) of 2.0×10^{-5} M. The fluorescence emission of the Zn(II) complex is bright red, whereas weaker intensity emission of the free chemosensor is hardly visible since it is at the red end of the visible spectrum, extending to near IR. Not surprisingly, the dipicolylamine ligand imparts remarkable metal ion selectivity onto this chemosensor (Figure 2). The chemosensor **4** generates the largest signal

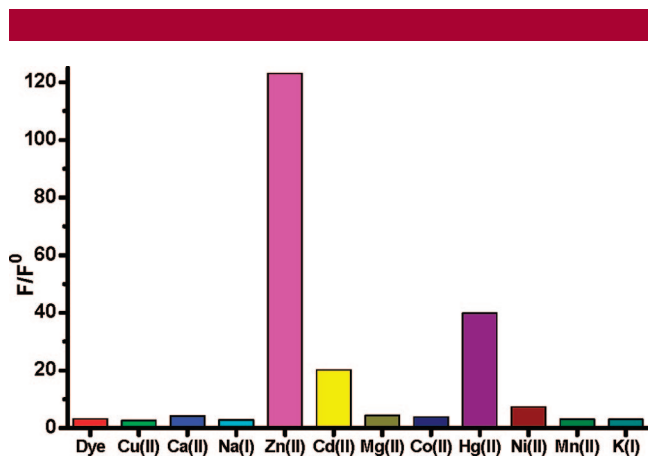


Figure 2. Normalized emission intensities at 680 nm (the emission intensity of the free dye = 1) of the chemosensor ($1.2 \mu\text{M}$) in the presence of selected metal ions (1.0 mM). Excitation was done at 630 nm with 5 nm slit widths. Solvent: ethanol–aqueous buffer mixture (5% in ethanol, HEPES 0.1 M, pH = 7.2).

in the presence of Zn(II) ions, but Hg(II) and Cd(II) also show some response. Fortunately, in biological media (unlike environmental samples), these three ions rarely compete, and our data show that most likely biological interferants such as Ca(II), K(I), Na(I), and Mg(II) have no effect on the emission spectrum. The results of a competition experiment between Zn(II) and selected metal ions is shown in Figure 3. Digital photographs (Figure 4) of the aqueous solutions of the chemosensor with added Zn(II) clearly show very large enhancement of the red emission under UV illumination at 360 nm, with no perceivable changes in the presence of other biologically relevant ions.

In conclusion, we have demonstrated that versatile bodipy chemistry allows the synthesis of long wavelength emitting, water-soluble dyes through stepwise reaction of the methyl groups at the 3 and 5 positions of the bodipy core. Thus, multiple and distinct functional groups can easily be placed on the chromophore, offering many opportunities for the

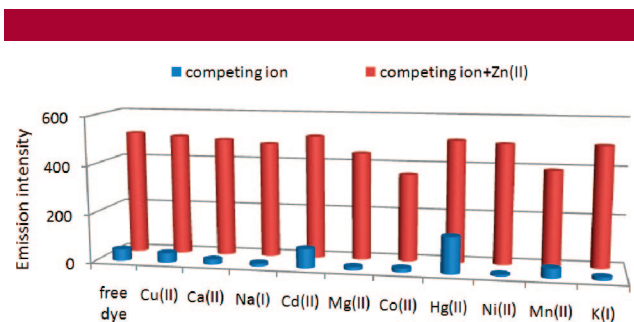


Figure 3. Results of the competition experiments between Zn(II) and selected metal ions. The free dye **4** (chemosensor) concentration was set at $5.0 \mu\text{M}$. Excitation was at 630 nm; emission intensity values at 680 nm were collected; and all metal ions were added at $200 \mu\text{M}$ concentration. Solvent: ethanol–aqueous buffer mixture (5% in ethanol, HEPES 0.1 M, pH = 7.2).

rational design for novel ratiometric chemosensors. The chemosensor discussed in this communication is a prototype



Figure 4. Digital photographs of the chemosensor solutions ($5.0 \mu\text{M}$) in the presence of different metal ions at $100 \mu\text{M}$ concentration. The upper plate is taken under ambient light, and the bottom one under UV illumination at 360 nm. Solvent: ethanol–aqueous buffer mixture (5% in ethanol, HEPES 0.1 M, pH = 7.2).

for such molecular ion sensors to be developed via this approach.

Acknowledgment. This work was supported by the Turkish Scientific and Technical Research Council (TUBITAK) and the Turkish Academy of Sciences (TUBA). The authors gratefully acknowledge Mr. Ilker Kutuk (METU) for his assistance with the synthesis work.

Supporting Information Available: Syntheses, experimental details, and additional spectroscopic data for all new compounds. This material is available free of charge via the Internet at <http://pubs.acs.org>.

OL801554T