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Feasibility of Modifying Existing Chemistry Demonstrations by Using Substitute Materials

Dane Mykel M. Quiambao^{1*}, Armando M. Guidote Jr.^{1,3}, Rhodora F. Nicdao²

¹*Department of Chemistry, Loyola Schools, Ateneo de Manila University, Katipunan Ave. Loyola Heights, Quezon City, Philippines 1108*

²*Department of Education, Loyola Schools, Ateneo de Manila University, Katipunan Ave. Loyola Heights, Quezon City, Philippines 1108*

³*Philippine Institute of Pure and Applied Chemistry (PIPAC), Ateneo de Manila University Campus, Katipunan Ave. Loyola Heights, Quezon City, Philippines 1108*

*Author to whom correspondence should be addressed; email: qdanemykel@gmail.com

ABSTRACT

This study aimed to reiterate the use of Chemistry demonstrations as effective teaching tools to students while addressing some of its drawbacks, which discourages teachers from doing them such as cost and safety. Four chosen existing Chemistry demonstrations (Blue Bottle Experiment, Copper Sulfate Experiment, Blown Away, Dancing Flames) were modified by using substitute reagents, which are more accessible, relatively safer, and at lower cost. These demonstrations were chosen based on how easily the substitute reagents will be obtained. Afterwards, with the permission of a private junior high school, they were presented to a group of Grade 9 students of their choosing. The students were asked to evaluate each demonstration using a Likert scale-based questionnaire. This questionnaire rates each demonstration in terms of aesthetics, the materials and procedure done, effectiveness to explain certain Chemistry topics, safety, and the students' overall judgment regarding the use of demonstrations as teaching tools. By converting their evaluation to quantitative values, the demonstrations scored high in all major categories. With this, it is highly recommended to explore other Chemistry demonstrations for possible modifications, which can be integrated in lecture classes.

Keywords: *Chemistry demonstrations; Introductory Chemistry; Likert scale-based questionnaire*

INTRODUCTION

In the field of Science, Chemistry has always been one of the more exciting sciences to learn about because different products and phenomena that we see in real life can be explained. It is also where the other branches of Science such as Biology and Physics come together, which makes it all the more interesting.

But at the same time, this makes it one of the more complicated sciences both to learn and to teach. As a student, one of the most challenging parts of learning Chemistry is being able to understand the concepts from the macro level. The fact that Chemistry primarily deals with particles that cannot be seen by the naked eye and that there are a lot of things going on makes it even more difficult. This is why Chemistry courses, or Science in general, should be supplemented by laboratory classes (Hered, 1950). But due to lack of facilities and equipment especially in public schools in developing countries, most students are not able to experience that. This is more difficult to teachers too because with the lack of lab classes, they resort to teaching by the book and rely on visual aids such as pictures and videos. What is worse is that some teachers cannot afford visual aids due to the lack of facilities and financial problems. This might decrease the students' interest in the subject and only a few or even none would want to pursue a career in the field in the future.

One of the alternatives to laboratory classes that can be done is by doing demonstrations wherein instead of just watching an experiment during a lecture, the teacher executes the experiment while the rest of the class observes. One of the skills, which they can still develop through this even without the typical laboratory experiments is problem-solving skills (Meyer et al., 2003). They will be encouraged to ask questions about what happened and even draw inferences of what they think happened during the demonstration. Their observational skills will also be honed just like in a typical laboratory setting ("Chemical Demonstrations in the Classroom", n.d.) wherein they will learn to be conscious of the different changes happening in the demonstration. This will also result in a constant student-teacher interaction as compared with a typical lecture wherein the professor teaches in almost a monologue. "Teaching methods that involve students and teachers are proven to be better," (Zejnilagić-Hajrić & Nuic, 2015) because they work together to make the learning experience more effective. A method of teaching that encourages team learning and student-led discussions (such as chemistry demonstrations), as compared to a normal lecture teaching, favors better performance of the students (Carpenter, 2006) since it encourages more participation in class boosting their self-confidence.

Despite this, using this method is still not widely used due to various reasons. One reason is because it is time-consuming (Walton, 2002) considering that the bulk of their time is allotted for preparation of their lesson plans (Meyer et al., 2003). Time is not only needed in preparing the demonstration, but it is also needed in trying-out the demonstration, preparing the necessary points that the educator wants to emphasize, and looking out for safety precautions that needs to be observed. Another hindrance is the misconception that demonstrations need expensive kits or materials (Meyer et al., 2003).

This study aims to modify pre-existing Chemistry demonstrations that can address the reasons stated above by using substitute reagents, which are better than the original in terms of accessibility, safety, and cost. Necessary revisions to the procedures were done to replicate the result of each original demonstration. Afterwards, as a supplement to the study, these demonstrations were presented to a group of high school students from a private junior high school. Each demonstration was evaluated using a Likert scale-based questionnaire which is based on the given criteria.

The use of demonstrations in lectures has been supported by a lot of studies worldwide over the past years (Walton, 2002; Meyer et al., 2003; Sullivan, 1990; O' Brien, 1991). A study conducted by Paul Walton, a chemistry faculty member at the University of York involved getting responses from freshmen undergraduate students who attended an Acids and Bases lecture with integrated demonstrations. They were given a set of statements and were asked if they agree to each or not. The results showed that 87% (either totally or partially) agreed that demonstrations helped them

understand the theories behind the lesson. Moreover, 95% (either totally or partially) agreed that the demonstrations kept their interest during the lecture (Walton, 2002).

Another study was conducted by Ophardt, Applebee, and Losey in 2005, which involved students from nonscience major courses in Elmhurst College. In their laboratory course, they performed various chemistry demonstrations in class and eventually in front of students from a local elementary school. The students were then asked to give their thoughts on the course in terms of their learning objectives and comparing it with the traditional laboratory setting. Their questionnaires have statements, which they rate in a 1-5 range, with 5 being the highest. The results showed high average scores when talking about their learning objectives. These include learning basic chemistry concepts (score: 4.67) and strengthening the interest in science (score: 4.28). Comparing with traditional science labs, students preferred their demonstration-focused lab. Areas of question include interest in activities performed on a given day (score: 4.42) and freedom to learn at their own pace (score: 4.00) (Ophardt, et.al., 2005). The latter result proves that chemistry demonstrations can help educators and adapt to the students' different learning styles (Meyer et al., 2003).

METHODS

Existing chemistry demonstrations were modified by replacing the reagents with more accessible materials with lower cost. The chemistry demonstrations were chosen based on the availability, accessibility, and affordability of possible substitutes, particularly in the Philippines. Since all the demonstrations are adequately documented in various sources, these became the visual models of what the demonstrations (experiments) should look like. They became the basis in possible changes in formulation of the materials, comparing the original and the substitute.

As a supplement to the study, the demonstrations were presented to a group of 24 junior high school students in a private junior high school. The sample size and respondents were chosen by the school itself based on their schedule as well as the relevance of the prepared demonstrations to their current curriculum. Afterwards, they were given a Likert scale-based questionnaire where they were asked to evaluate the presented demonstration (See Supporting Information A). This work was granted a Validation of Exemption from Review by the Ateneo de Manila University Research Ethics Office (UREO).

The 10-item questionnaire is based on five major categories that was deemed important in evaluating a Chemistry demonstration. Items # 1 and 6 fall under the Attention Getting category, which rates how the demonstration got the students' attention visually. The second category is Technical Procedure (items # 2 and 4), which assesses the demonstration materials and procedure. This also includes accessibility and difficulty. The third category is Lecture Value (items # 3, 5 and 7), which deals with the effectiveness of the demonstrations to the respective Chemistry topics that they were connected to. Fourth category is Safety (item # 8), which rates the overall safety of each demonstration, i.e., if it is possible to do each demonstration outside the laboratory and if it is without any extreme danger to the students and teacher. The last category is Overall Judgment (item # 9 and 10), which rates the overall reaction of the students to the demonstrations. It rates how helpful the demonstrations can be if they are applied as teaching tools during lectures.

The following are the modified demonstrations:

1. Blue Bottle Experiment – this involves the reduction-oxidation reaction of dextrose and methylene blue in a basic solution (See Supporting Information B);
2. Copper Sulfate Experiment – this is a reversible reaction which involves the dehydration of copper (II) sulfate (See Supporting Information C);
3. Blown Away – this is a precipitation reaction between calcium hydroxide and carbon dioxide (See Supporting Information D); and
4. Dancing Flames – this is a single displacement reaction between copper (II) sulfate and aluminum in acid (See Supporting Information E).

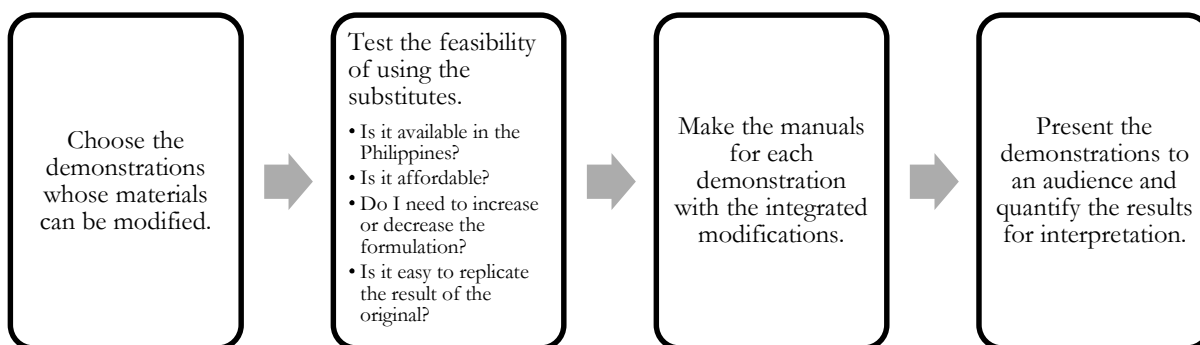


Figure 1. Procedural Diagram for the Study

RESULTS AND DISCUSSION

The main reagents of each demonstration as written and published in online journals and respected websites are shown in Table 1 indicating as well the list of the substitute materials and their cost of purchase at the time of the study. Using these substitutes, the results of the original Chemistry demonstrations were replicated with some modifications in the procedure.

The Likert scale responses were translated into a numerical value using the criteria suggested by Joshi et al. (2015) where the Likert scale can be treated as an interval scale, which is usually a five or seven-point scale (“Likert Scale | Simply Psychology,” n.d.). The mean score of each major category of each experiment was obtained. These are the ‘interval estimates’ (Joshi et al., 2015), which were then fit to a score range. Each were given a corresponding interpretation to obtain a qualitative result. Afterwards, these mean scores were combined to report all demonstrations as one in each of the categories.

Table 2 showed that all the presented Chemistry demonstrations afforded positive results in all categories, with the mean scores equivalent to the Very Good to Excellent range. Supported by the low standard deviations measured from all categories, it can be inferred that the students enjoyed the demonstrations and that they saw the demonstrations being beneficial in understanding the Chemistry topics.

Table 1. Main reagent list of the experiments and its substitute materials

Experiment	Original Experiment Materials		Substitute Materials	
	Source: Royal Society of Chemistry ("The 'Blue Bottle' Experiment", n.d.)	Source: Flinn Scientific ("Blue Bottle Experiment", n.d.)	Substitute	Price of Substitute
I. Blue Bottle Experiment	Potassium hydroxide, solid	Potassium hydroxide, solid	Liquid sosa	PHP 92.50 (USD 1.86) (500mL)
	Glucose (dextrose), solid	Dextrose, solid	Dextrose powder (dog supplement)	PHP 120.00 (USD 2.41) (300g)
	Methylene blue, solid	1%aq methylene blue solution	Methylene blue solution (fish tank disinfectant)	PHP 25.00 (USD 0.50) (120mL)
	Ethanol	~	~	
II. Copper Sulfate Experiment	Source: Royal Society of Chemistry ("A Reversible Reaction of Hydrated Copper (II) Sulfate", n.d.)	Source: UC San Diego ("Demo 51: Dehydration and Rehydration of Copper Sulfate", n.d.)	Substitute	Price of Substitute
	Copper (II) sulfate (VI)-5-water solution	Copper (II) sulfate, solid, lab grade	Swimming pool algaecide powder (chelated copper sulfate)	PHP 550.00 (USD 11.05) (1kg)
III. Blown Away	Source: Education ("The Limewater Carbon Dioxide Test, Science Project, Education.Com", n.d.)	Source: Microscale Gas Chemistry (Mattson, 2018)	Substitute	Price of Substitute
	Calcium oxide in distilled water	Limewater	Pickling lime (food grade calcium hydroxide) in tap water	PHP 100.00 (USD 2.01) (500g)
IV. Dancing Flames	Source: Royal Society of Chemistry (Fleming, n.d.)	Source: Flinn Scientific ("Foiled" Again, n.d.)	Substitute	Price of Substitute
	Copper (II) chloride * dihydrate, solid	Copper (II) sulfate solution	Swimming pool algaecide powder (chelated copper sulfate)	PHP 550.00 (USD 11.05) (1kg)
	1M hydrochloric acid	~	28% muriatic acid	PHP 49.50 (USD 0.99) (100mL)
	~	1M sodium chloride solution	Table salt	

*At an exchange rate of PHP 1.00 = USD 0.020

Table 2. Summary of Derived Data

Summary (All Demonstrations)	Results			Reference scale	
	Mean	Standard Deviation	Interpretation	Interpretation	Score Range
Attention Getting (1,6)	6.25	0.415	Excellent	Excellent	6.15 – 7.00
Technical Procedures (2,4)	6.14	0.279	Very Good	Very Good	5.30 – 6.14
Lecture Value (3,5,7)	5.73	0.151	Very Good	Good	4.44 – 5.29
Safety (8)	6.05	0.176	Very Good	Satisfactory	3.58 – 4.43
Overall Judgment (9,10)	6.68	0.439	Excellent	Fair	2.72 – 3.57
				Poor	1.87 – 2.71
				Very Poor	1.00 – 1.86

These results support the studies regarding Chemistry demonstrations as helping to “engage students with chemistry itself,” (Meyer et al., 2003). For instance, Attention Getting has a mean score of 6.25. The students thus agree that the use of Chemistry demonstrations in general draws the students’ attention, which helps build their interest in the topic. This result is similar to Walton’s study, which stated that 95% of the surveyed students said that demonstrations kept their interest in the lecture (Walton, 2002). On the other hand, Technical Procedures (score: 6.14) and Safety (score: 6.05) having high mean scores can address the hesitations the educators have in doing demonstrations as highlighted in this study. Addressing the safety and accessibility concerns by using cheaper and safer substitute reagents suggests strong potential of these experiments for easy integration to lesson plans and curricula. The feasibility of the modified experiments using substitute materials have been successfully demonstrated to achieve its purpose in reinforcing a topic in Chemistry.

CONCLUSION

The study was successful in modifying existing Chemistry demonstrations by using substitutes that are more accessible, relatively safer, and at a lower cost. With some modifications in the procedure, the end result of each original demonstration was replicated. The Chemistry demonstrations were evaluated with positive results in all categories. This shows that the modified selected demonstrations can be used with confidence. The positive feedback also addresses, and possibly removes, some underlying reservations such as cost and safety in allowing students to do the experiments. In general, this study can help further push and encourage the use and integration of Chemistry demonstrations in school curricula today.

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SUPPORTING INFORMATION A: Likert Scale

As part of the reported results of the current study, the principal investigator would like you to answer this questionnaire about the chemistry demonstration that you have watched. Please answer as objectively as you can. Rest assured that only the principal investigator will see your individual responses and you will not be asked for any personal information.

Please check the box which describes your response to the following observations.

Note: The term “chemistry demonstration” refers to the experiment presented, not the whole session.

The chemistry demonstration...	Strongly Disagree	Disagree	Slightly Disagree	Neutral	Slightly Agree	Agree	Strongly Agree
.. was aesthetically pleasing.							
.. was easy to do and the materials were accessible.							
.. was in line with the topic/s that it was discussed into.							
.. was poorly executed and hard to follow.							
.. made the related topic more confusing.							
.. was boring and did not catch my attention.							
.. helped me understand the related topic better.							
.. is safe to do outside the lab and its safety precautions were laid out well.							
Chemistry demonstrations are helpful tools and should be integrated in lectures.							
Presenting Chemistry demonstrations make the lecture more exciting and interesting.							

Other comments/suggestions:

SUPPORTING INFORMATION B Experiment 1: Blue Bottle Experiment

Related topic/s: Redox Reactions

Objective: to visualize an example of a redox reaction through color change

Materials

5g (~1 ½ tsp) dextrose powder (food supplement for dogs)

3mL (~1 tsp) liquid sosa (NaOH solution)

5 to 7 drops methylene blue (sold in pet shops)

500mL plastic bottle

Safety

- Use gloves in handling the NaOH solution and methylene blue.
- Make sure that the plastic bottle used does not have any holes.

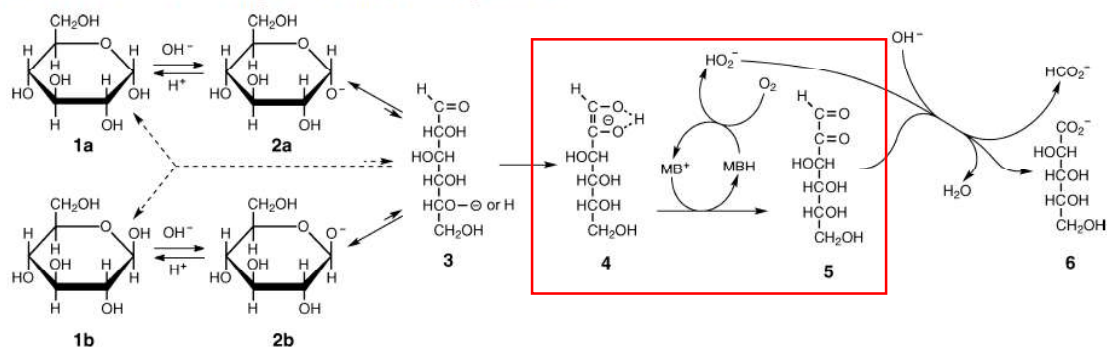
Procedure

1. Dissolve 5g dextrose powder in 300mL water in the plastic bottle.
2. Add 3mL NaOH solution and shake.
3. Mix in 1mL methylene blue and leave to settle.

Explanation

Different reactions are taking place during the color change as seen below but the more important reaction is the reduction and oxidation of methylene blue. When the bottle turns blue to colorless, methylene blue is reduced by the dextrose powder through its enediolate anion in alkaline solution to produce methylene white (MBH) which is another name for reduced methylene blue. On the other hand, when the bottle turns colorless to blue, methylene white (MBH) is oxidized by oxygen inside the bottle (See 4-5 in the figure) (Anderson et al., 2012).

Scheme 1. Reactions Occurring in the Blue Bottle Experiment^a



^a 1a, α -D-glucopyranose; 1b, β -D-glucopyranose; 2a, α -D-glucopyranosyloxy anion; 2b, β -D-glucopyranosyloxy anion; 3, open chain form of glucose; 4, enediolate anion; 5, D-arabino-hexos-2-ulose (glycosulose, open chain form); 6, D-arabinonate anion; MB⁺, methylene blue; MBH, methylene white; HO₂⁻, hydroperoxide anion; HCO₂⁻, formate anion.

To put it simply,

In prep: glucose + MB⁺(blue) → glucosone + MBH (clear)

In shaking: MBH (clear) + O₂(g) → MB⁺ (blue) + HO₂⁻

After oxygen exhaustion: MB⁺ (blue) → MBH (clear)

Disposal

Since all reagents used are diluted, the mixture can be poured down the sink in running water. Rinse the plastic bottle first before throwing it in the recycle bin.

Documentation



Materials used



Blue Bottle in idle

Reference

1. Anderson L, Wittkopp SM, Painter CJ, Liegel JJ, Schreiner R, Bell JA, et al. What Is Happening When the Blue Bottle Bleaches: An Investigation of the Methylene Blue-Catalyzed Air Oxidation of Glucose. *J Chem Educ.* 2012 Oct 9;89(11):1425–31.

SUPPORTING INFORMATION C

Experiment 2: Copper Sulfate Experiment

Related topic/s: Reversible Reactions

Objective: to visualize the reversible activity of a chemical reaction

Materials

- Chelated copper sulfate/pool algacide (sold in hardware stores)
- Test tube and test tube holder / spoon
- Medicine dropper
- Water
- Makeshift burner
 - 3 empty soda cans
 - Iron nail and hammer
 - Scissors
 - Small knife
 - Long nose pliers
 - Ethyl alcohol
 - Tongs

Safety

- Be careful in using the knife.
- In heating up with the test tube, make sure that its mouth is away from you and the audience.
- Always keep your eyes on the burner when it is lighted. Make sure that no combustible materials are surrounding the burner which may catch fire.

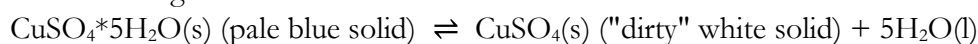
Procedure

- A. Making the makeshift burner (Lee Yu et al., 2015)
 1. Cut two of the three cans two inches from the base. These will serve as the base and cover of the burner.
 2. Cut the ends of the third can, leaving about three inches of the body, and cut along its height. This will serve as the support.
 3. With one of the two cans that are cut two inches from the base, trace the inner edges of its base with a knife. Puncture a hole using a nail and use the long nose pliers to cut through the base and pull it out.
 4. Staple the ends of the support to form a cylinder with a diameter equal to the rim of the base.
 5. Place the cover upside down over the base. Puncture pinholes along the top frame using a small nail.
 6. To light it up, half-fill the burner with ethyl alcohol and use matches. To be safe, use tongs in pointing the lighted matchstick to the solution.
- B. Demonstration
 1. Add chelated copper sulfate in the test tube.
 2. Using the test tube holder, hold the test tube at the burner. Make sure that the mouth of the test tube is away from you. Observe what happens.

3. Afterwards, wait for a few seconds for the powder to cool down.
 4. Again, with the mouth of the test tube away from you and the audience, slowly put several drops of water but not too much to avoid making it a mixture. Observe what happens.
 5. You may also use a spoon instead of a test tube to hold the chelated copper sulfate.
- Note: The time it takes to change its color completely depends on the amount of copper sulfate used.

Explanation

The following reversible reaction is as follows:



Because it is a reversible reaction, we can see that we are able to visualize the original reactants and the direction of the reaction. In the experiment, heating up the chelated copper sulfate colored blue evaporates the water thus making it colored white. On the other hand, adding drops of water to the heated copper sulfate changes its color back to blue (Nuffield Foundation and Royal Society of Chemistry).

Disposal

- Since the copper sulfate is still primarily solid, you may dispose it in the trash bin. Otherwise, you may dissolve it in water and pour in the sink. For the other materials, you may throw them in the trash bin.
- For the burner, make sure that the fire has been put out before throwing it away. You may do this by covering the burner. The burner can be reused so you may want to keep it.

Documentation



Materials used



Gradual discoloration of copper(II) sulfate

References

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SUPPORTING INFORMATION D

Experiment 3: Blown Away

Related topic/s: Formation of precipitates, single displacement reactions

Objectives:

- To demonstrate one of the main features of a chemical change which is the formation of a precipitate
- To visualize an example of a double displacement reaction

Materials

- 3g (~1tsp) pickling lime (food grade calcium hydroxide)
- Small cup (plastic or glass)
- Drinking straw
- Small spoon (for scooping the calcium hydroxide)
- 500ml plastic bottle

Safety

- Do not touch the calcium hydroxide with your bare hands. Use a spoon to scoop. To be safer, use gloves when doing the demonstration.
- Be careful not to suck in the liquid with the straw. Since it is food grade calcium hydroxide, it is generally safe when ingested at a small quantity.

Procedure

A. Before demonstration

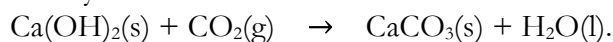
1. Dissolve 3g calcium hydroxide/pickling lime in water.
2. Pour to a 500ml plastic bottle and fill. Let it set for 24 hours.
3. Take the liquid. This is what you will use in the demonstration.

B. Demonstration Proper

1. Put some of the liquid in a glass. Blow into it using the straw until a change occurs. Observe.

Explanation

This experiment involves the double displacement reaction of calcium hydroxide and carbon dioxide from your breath:



As you blow into the solution, more calcium carbonate is produced which is causing the solution to cloud up and eventually turn white.

Disposal

- The end solution may be thrown in the sink with running water. The rest of the materials may be thrown in the trash bin.

Documentation



Materials used



Precipitation of calcium carbonate

Reference

1. The Limewater Carbon Dioxide Test | Science project | Education.com [Internet]. [cited 2019 Mar 13]. Available from: <https://www.education.com/science-fair/article/gas-sniffers/>

SUPPORTING INFORMATION E

Experiment 4: Dancing Flames

Related topic/s: Reactivity series, single displacement reactions, reduction-oxidation reactions, electrochemical reactions

Objectives:

- To demonstrate an example of a single displacement reaction
- To demonstrate the reactivity of aluminum

Materials

- 5g (~1 tsp) copper (II) sulfate/chelated copper sulfate-based pool algicide
- 50mL (~10 tsp) muriatic acid
- 1 250mL Erlenmeyer flask (or any heat-resistant container with a long neck)
- Table salt
- Water
- Mixing containers
- Matches/lighter

Safety

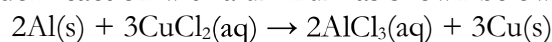
- Be careful in lighting up the solution. Make sure that no combustible materials are surrounding the setup which may catch fire. Use goggles for safety.
- When pointing the fire at the flask, make sure that you and your audience are not too close.
- For safety, use gloves especially when dealing with muriatic acid. If it spills on your skin, wash affected area with running water.

Procedure

1. Cut aluminum foil about 20cm long and a width as long as the base of the flask. Roll it loosely and put it inside the flask. Make sure that it is sitting on its side at the base of the flask. You can use the splint to push it down almost flat in the flask.
2. Dissolve 1 teaspoon (~5g) table salt in 50mL water. Set aside.
3. Dissolve the copper (II) sulfate in muriatic acid. Add half of the NaCl solution and 2 teaspoons (~10mL) water.
4. Pour this solution onto the flask with the aluminum foil. Be ready to light up the match or lighter.
5. Point the fire at the mouth of the flask as soon as you see gas emerging. You should hear a small pop and see a green flame “dancing” inside the flask.

Explanation

This experiment is derived from an earlier version, wherein copper(II) chloride was used to give a single replacement redox reaction with aluminum as shown below (“Foiled Again”):



Looking at the reactivity series, we see that aluminum is more reactive than copper thus the single replacement reaction occurs.

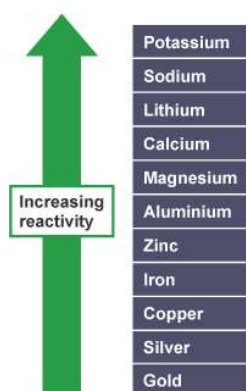
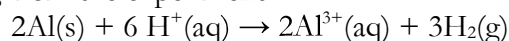
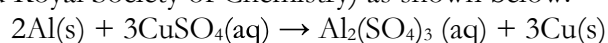


Figure 1. Reactivity Series (“f) Reactivity Series”, 2017)

At the same time, a side reaction occurs wherein free hydrogen ions from the copper(II) chloride solution, which is slightly acidic, react with aluminum to form hydrogen gas which is the same gas as the one emerging from the experiment.



In this experiment, we used copper(II) sulfate instead of copper(II) chloride because copper(II) sulfate is more accessible and less costly to buy. Using this alone will not push the reaction, which is why a solution of sodium chloride was used. This is because aluminum foil has a protective oxide layer which protects the aluminum metal from further reaction with air, water or acid (Fleming). Chloride ions disrupt this layer thus allowing the aluminum-copper reaction (Nuffield Foundation and Royal Society of Chemistry) as shown below:



The side reaction mentioned above is also applicable. Hydrogen gas is a product in the reaction, which is proven by the small pop heard when lighted.

Disposal

- The solution may be thrown in the sink with running water. Any unreacted aluminum foil may be thrown in the trash bin.

Documentation



Materials used



Green flame on top of dissolving aluminum

References

1. Foiled Again [Internet]. [cited 2020 Feb 5]. Available from: <https://www.flinnsci.com/foiled-again/dc95000/>
2. f) Reactivity Series [Internet]. IGCSE Chemistry revision. 2017 [cited 2020 Feb 5]. Available from: <https://igcsechemrevision.wordpress.com/section-2/f-reactivity-series/>
3. Fleming D. Dancing flames [Internet]. RSC Education. [cited 2020 Feb 5]. Available from: <https://edu.rsc.org/exhibition-chemistry/dancing-flames/2000045.article>
4. The real reactivity of aluminium [Internet]. RSC Education. [cited 2020 Feb 5]. Available from: <https://edu.rsc.org/resources/the-real-reactivity-of-aluminium/439.article>