Original scientific paper

Nickel contamination analysis at cost-effective silver printed paper-based electrodes based on carbon black dimethylglyoxime ink as electrode modifier

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Abstract

Electrochemical detection of metal cations at paper-based sensors has been suggested as an attractive alternative to current spectroscopic and chromatographic detection techniques due to the ease of fabrication, disposable nature, and low cost. Herein, a novel carbon black (CB), dimethylglyoxime (DMG) ink is designed as an electrode modifier in conjunction with 3-electrode inkjet-printed paper substrates for use in the adsorptive stripping voltammetric electroanalysis of nickel cations in water samples. The developed method provides a novel, low-cost, rapid, and portable adsorptive stripping detection approach towards metal analysis in the absence of the commonly used toxic metallic films. The study demonstrated a novel approach to nickel detection at paper-based sensors and builds on previous work in the field of paper-based metal analysis by limiting the use of toxic metal films. The device sensitivity is improved by increasing the active surface area, electron transfer kinetics, and catalytic effects associated with non-conductive dimethylglyoxime films through CB nanoparticles for the first time and confirmed by electroanalysis. The first use of the CB-DMG ink allows for the selective preconcentration of analyte at the electrode surface without the use of toxic Mercury or Bismuth metallic films. Compared to similarly reported paper-based sensors, improved limits of detection (48 μg L^{-1}), selectivity, and intermetallic interferences were achieved. The method was applied to the detection of nickel in water samples well below World Health Organization (WHO) standards.

Keywords

Carbon black; dimethylglyoxime; nickel; stripping voltammetry

Introduction

Remediation of metal pollution and its quantitative analysis remains a major area of focus for analytical chemists worldwide with considerable funding and time dedicated to novel and innovative means for improving current methods. Heavy metal pollution continues to garner significant attention due to its hazardous effects on the environment and organisms owing to its non-biodegradable and toxic nature [1]. Current spectroscopic and chromatographic techniques for monitoring this pollution are limited by their high costs, cumbersome equipment, and skilled operation [2]. Therefore, much attention has been placed on low-cost, rapid, and portable means of analysis. Electrochemical detection of metal ions is one such technique that offers highly selective, repeatable, and sensitive quantitative analysis [3]. To date, a range of metal pollutants has been determined by electroanalysis including heavy, platinum group, and transition metals. Nickel determination has been studied following several approaches, however, adsorptive stripping voltammetry remains the most effective technique utilized [4]. The method relies on the reduction of the metallic species from a stable metal-centered coordination complex formed with inorganic chelating agents or biological materials [4,5]. Several metallic films have been suggested to aid in the accumulation of the metal analyte at the electrode surface [6-8], but work has been done to improve the detection capability when a metallic film is not employed.

Carbon nanomaterials have been extensively studied as electrode modifiers in electrochemical sensor applications. Graphene [9,10] carbon nanotubes [11,12], carbon dots [13,14], fullerenes [15], carbon onions, and other nano-sized carbon structures have been effective in improving the sensitivity of sensors owing to their high active surface, large surface-to-volume ratio, improved electron transfer properties, increased stability, and wide active potential window. These nanoscale carbon sensors have been utilized in a diverse array of applications, however, the cost of synthesis, tedious preparation methods, and lack of availability hinder their use. To combat this, carbon black (CB) nanoparticles have been proposed as electrode modifiers for cannabis seed oil, antihypertensive drugs, phenolic compounds, DNA, flavonoids, and metal analysis. Carbon black offers comparable results to other carbon nanostructured sensors[16-23] It is a commonly available carbon source with elemental carbon arranged in aggregates of fine particles from 3 - 100 nm. The nanomaterial is often employed as thin film or electrode coating in conjunction with other nanoparticles and active materials. Recently, few studies have proposed the use of CB-based sensors to enhance the sensitivity of electrochemical sensors for stripping voltammetric analysis [24-28]. However, few studies have been conducted for metal analysis. To date, no work has been reported on the use of carbon black in conjunction with chelating agents to form composite materials to improve detection sensitivity and selectivity in adsorptive stripping voltammetric detection of metal ions.

Paper-based sensor use has steadily gained traction as an analytical technique in recent times due to the increased availability, ease of modification, capillary flow, and biodegradable nature [29]. Many standard analytical systems have been adapted to paper-based methods including colorimetric and electrochemical detection. The excellent sorption behaviour of the paper-based devices makes it an attractive option for these analysis systems creating lateral flow and origami devices particularly for use in resource-limited settings on the African continent where access to costly instrumentation and skilled laboratories is lacking. Due to this access to proper medical care is delayed causing severe illnesses. Hydrophobic paper substrates have also been studied by several laboratories to replace costlier carbon, metallic and ceramic substrates through coating via printing techniques. While screen-printing is the simplest and most employed method, inkjet printing has been suggested as a suitable alternative [30-32]. Owing to these benefits, electrochemical sensors

[33-38] and energy devices [39-43] have recently been reported based on inkjet printing. Inkjet-printed paper electrodes have previously been studied by our research group where a graphene, gold nanoparticle composite was studied in the presence of a mercury film for Nickel detection. To date, a number of works have reported the use of paper-based sensors for stripping voltammetric analysis of metal ions. Here, screen-printed electrodes, paper disks, and microfluidic devices have been studied. However, only the work by Pokpas *et al.* has shown the applicability of paper-based devices to be used in adsorptive stripping voltammetry for metal ions. This study demonstrated the first use of nanostructured carbon materials for use in adsorptive detection mechanisms for metal analysis. The low sensitivity of paper-based sensors remains the biggest drawback of the technique to date [7,44]. As a result, metallic films like mercury are still readily employed. Further work to improve the sensitivity for detection towards metal cations is required before commercial use of these is implemented and particularly to eliminate the use of commonly used amalgam and alloy formation. In addition, no work has been reported to utilize carbon black inks over costlier nanomaterials to enhance the sensitivity of paper-based electrodes.

In this work, a disposable inkjet-printed paper electrode is fabricated for the rapid, cost-effective, and portable detection of Nickel, metal cations in water samples as an electrochemical sensor. The detection sensitivity and selectivity of the low-cost silver printed paper electrodes were improved through modification with carbon black dimethylglyoxime (CB-DMG) paste as electrode modifier in the absence of a metallic film. The CB-DMG allows for effective preconcentration of the electrode surface with Nickel cations in the form of Ni-DMG complexes while maintaining excellent electron transfer properties required for electrochemical detection. The study builds on previous work by our research group for use of inkjet-printed paper-electrodes in adsorptive stripping voltammetric detection of metal ions. Furthermore, the study showcases an attractive alternative to the use of toxic mercury films previously studied. The developed, point-of-care sensor was applied to the detection of metal ions in tap water samples below the World health organization (WHO) limits of detection of 0.1 mg L^{-1} and demonstrates for the first time a carbon black ink combined with a chelating agent, dimethylglyoxime as electrode modifier.

Experimental

Apparatus

A Fujifilm Dimatix DMP-3281 materials deposition printer was employed for all inkjet printing following DesignCAD templates onto Penguin photo paper. All electrochemical voltammetric experiments were performed with a Metrohm Autolab PGSTAT101 instrument, in combination with the Nova 1.11 software, and controlled by a personal computer. All electrochemical experiments were performed in one compartment 20 mL voltammetric cells at room temperature unless stated otherwise.

Chemicals and reagents

All chemicals used in the study were of analytical reagent grade. Ultra-pure distilled water (Millipore) was used to prepare all solutions. Harima NPS-JL-nano-silver ink, for low-temperature curing (Harima Chemicals Group) and carbon conductive paste for screen printing (Gwent Group), was used without modification. Nafion perfluorinated resin solution 5 wt.% in lower aliphatic alcohols and water, 2,3-Butanedione-dioxime (Dimethylglyoxime) and Nickel(II) standard stock solutions (1 g L^{-1}, atomic absorption standard solution) were obtained from Sigma-Aldrich and diluted as required. Carbon black was purchased from Alfa Aesar. Ammonia/Ammonium Chloride

(NH_3/NH_4Cl) buffer solution (0.1 M, pH 9.3), prepared by mixing appropriate quantities of ammonia (NH_3) and ammonium chloride (NH_4Cl) was used as supporting electrolyte for all studies. A Metrohm 827 pH lab pH meter was calibrated using pH 4 and 7 calibration buffer solutions and then used to verify the pH of the prepared NH_3/NH_4Cl buffer solution.

Design and fabrication of carbon-coated AgNP inkjet-printed electrodes

A similar procedure described in our previous work was utilized for inkjet-printed electrode preparation. Generally, a piezoelectric materials printer was used to carry out all inkjet printing of Harima NPS-JL-nano-silver ink onto hydrophobic photographic paper.

Preparation of nafion, dimethylglyoxime, carbon black, silver printed paper-based electrodes (N-DMG-CB-Ag-PPE)

Appropriate quantities of carbon black (0.8 g) and dimethylglyoxime (0.2 g) were mixed in a nafion solution (0.2 wt. %, in EtOH) to prepare the nafion-dimethylglyoxime-carbon black (N-DMG-CB) ink. The viscosity of the prepared ink was adjusted using EtOH if needed. A variety of CB: DMG ratios were prepared as needed. Consecutive, 5 μL aliquots of as-prepared N-DMG-CB ink were drop-cast and screen printed onto the working electrode to prepare the Nafion/Dimethylglyoxime/Carbon black/Silver Printed Paper Electrodes (N-DMG-CB-Ag-PPEs). N-DMG-CB-Ag-PPEs were dried at 85 $^{\circ}$C for 1 hour.

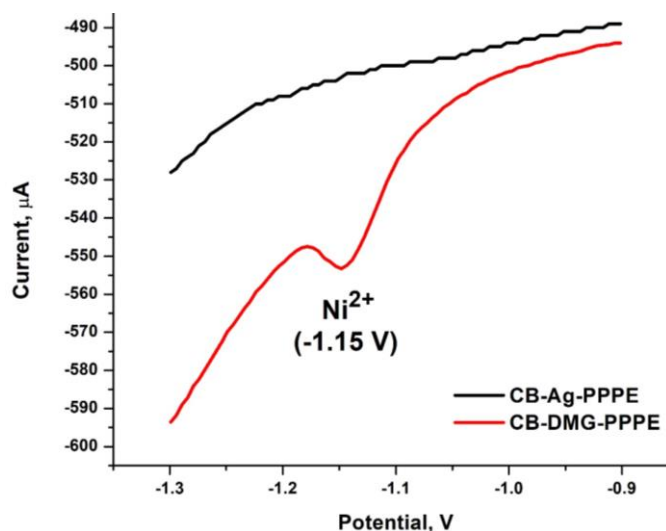*Procedure for the SW-AdCSV analysis of Ni^{2+}*

The N-DMG-CB-Ag-PPE was used to analyze a 10 mL sample of 0.1 M ammonia buffer (pH 9.4) Ni^{2+} containing 2 mM DMG and 10 mg L^{-1} Hg^{3+} respectively. The electrochemical detection was carried out by cathodic stripping voltammetry between -0.7 V and -1.35 V with a 120 s accumulation time at -0.7 V. f = 5 Hz and Amplitude = 10 mV.

Results and discusion

Ni^{2+} detection at the CB-DMG-Ag-PPPE

Silver printed working electrodes have shown low sensitivity towards [Ni(dmgH$_2$)] complex detection and a narrow active potential window, as discussed in previous sections. Carbon inks have been shown to significantly improve the active potential window and have shown a clear affinity towards the desired application. Carbon black ink, formed from the nanoparticle form of carbon was proposed as an ultrasensitive carbon coating material to further improve the electrode sensitivity of inkjet-printed paper-based sensors in the absence of an electroplated metallic film. Figure 1 represents typical square-wave adsorptive cathodic stripping voltammograms (SW-AdCSV) for 300 μg L^{-1} Ni^{2+} at the carbon black silver printed paper electrode (CB-Ag-PPPE) and carbon black/dimethylglyoxime silver printed paper electrode (CB-DMG-Ag-PPPE) respectively. The CB-Ag-PPPE shows no cathodic stripping peaks in the active potential window between -0.7 and -1.3 V in 0.1 M NH_3/NH_4Cl Buffer (pH 9.4) with deposition potential (- 0.7 V), deposition time (120 s), frequency (5 Hz), amplitude (0.1 V) and voltage step (0.005 V). In contrast, the CB-DMG-Ag-PPPE shows a sharp, well-resolved stripping peak at -1.14 V for Ni^{2+} reduction from the formed [Ni(dmgH$_2$)] adsorption complex at the electrode surface. Immobilization of the chelating agent at the electrode surface in conjunction with the high active surface area and conductive carbon nanoparticles provides sensitive and selective detection of Ni^{2+} cations. The overall electrolytic reaction may be described in the two-step process below (Equations 1 and 2):

Ni^{2+} + 2 dmgH$_2$ +2 OH$^-$ $\rightarrow$[Ni(dmgH$_2$)] + 2 H$_2$O Pre-concentration Step   (1)

[Ni(dmgH$_2$)] + 10 e$^-$ + 10 H$^+$ $\rightarrow$ Ni^{2+} + 2 DHAB Stripping/reduction step   (2)

Figure 1. *SW-AdCSV of 300 µg L^{-1} Ni^{2+}, obtained at the (a) CB-Ag-PPPE and (b) CB-DMG-Ag-PPPE. Supporting electrolyte (0.1 NH$_3$/NH$_4$Cl Buffer, pH 9.4), deposition time (90 s), deposition potential (- 0.7 V), rotation speed (1000 rpm), frequency (5 Hz), amplitude (0.1 V)*

Influence of chelating agent concentration in CB-DMG Ink

The detection of Ni^{2+} is highly dependent on the formation of the [Ni(dmgH$_2$)] and its subsequent adsorption onto the working electrode surface. A sensitive technique was established by immobilizing the chelating agent at the electrode surface, simplifying the stripping voltammetric mechanism. Here, its application at printed paper electrodes was achieved. Optimization of the dimethylglyoxime loading within the prepared carbon black ink (CB:DMG) is illustrated in Figure 2 below. The CB:DMG ratios were varied between 1:0, 1:4, 1:2, and 1:1, and its effect on the stripping voltammetric peak currents were investigated towards 300 µg L^{-1} Ni^{2+} detection by AdCSV. The absence of a chelating agent (CB:DMG, 1:0) prevents adequate pre-concentration of Ni^{2+} at the electrode surface and no cathodic stripping peak is observed, as previously stated.
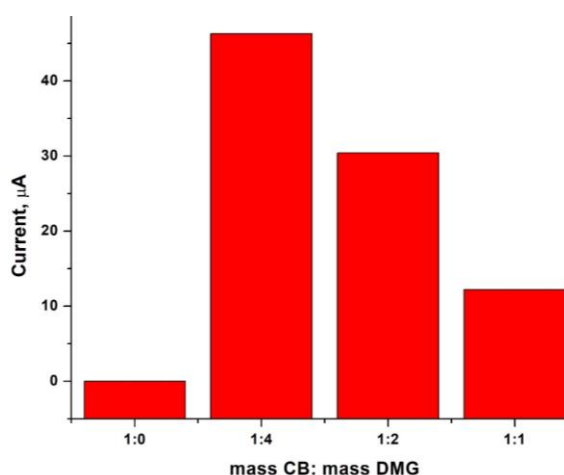
Figure 2. *Influence of DMG loading within carbon black inks on the stripping peak current of 300 µg L^{-1} Ni^{2+}.*

The addition of a DMG chelating agent at the working electrode facilitates the adsorption of metal cations onto the electrode surface and the subsequent formation of metallo-chelate complexes. Ni^{2+}, present in larger concentrations can now be reduced and a stripping peak current achieved. The inclusion of more non-conductive DMG inhibits electron flow and a decrease in the measured peak current is seen. A 1:4 weight ratio of CB:DMG yielded the optimum stripping peak currents and was therefore selected for all further experiments.

Morphological characterization of CB-DMG-Ag-PPPE

The morphological characterization of the prepared electrodes was studied by scanning electron microscopy. Figure 3 shows the high-resolution electron microscope (HRSEM) images of (a) unmodified inkjet printed AgNP electrode, (b) carbon black modified Ag-electrode and (c) dimethylglyoxime, carbon black Ag-electrode. The HRSEM image of the unmodified AgNP electrode shows a uniform film of spherical Ag nanoparticles deposited on the photographic paper substrate. An even coating with no distinct defects in the printing process is observed. Figure 3B demonstrates 'bushy' aggregates of carbon black particles deposited on the silver WE surface shown by spherical particles. Carbon black nanoparticles with an average particle size of ~ 80 nm are observed. Adequate surface coverage of the printed paper electrode surface is observed with a somewhat uniform distribution. Upon inclusion of dimethylglyoxime chelating agent in a 1:4 ratio, crystalline deposits can be seen within the carbon black film with an uneven surface distribution as indicated in Figure 3C and confirm the formation of the CB-DMG composited electrode.
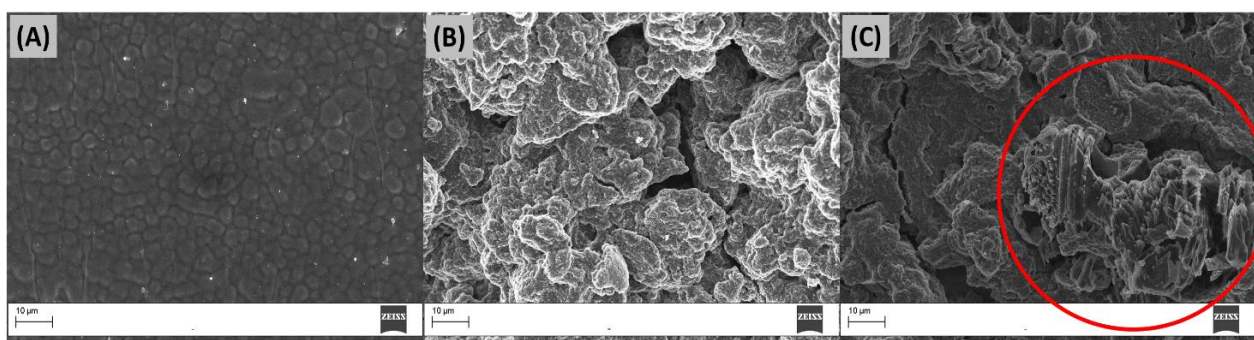

Figure 3: High resolution scanning electron microscope images of (a) inkjet printed AgNPs, (b) CB-AgNP and (c) DMG-CB-AgNP working electrodes on photographic paper at 1 kX magnification.

CB-DMG-Ag-PPPE instrumental parameter *optimization*

The square wave instrumental parameters affecting the analytical response of the CB-DMG-Ag-PPPE; namely deposition potential, deposition time, frequency, and amplitude were optimized and illustrated in Figure 4.
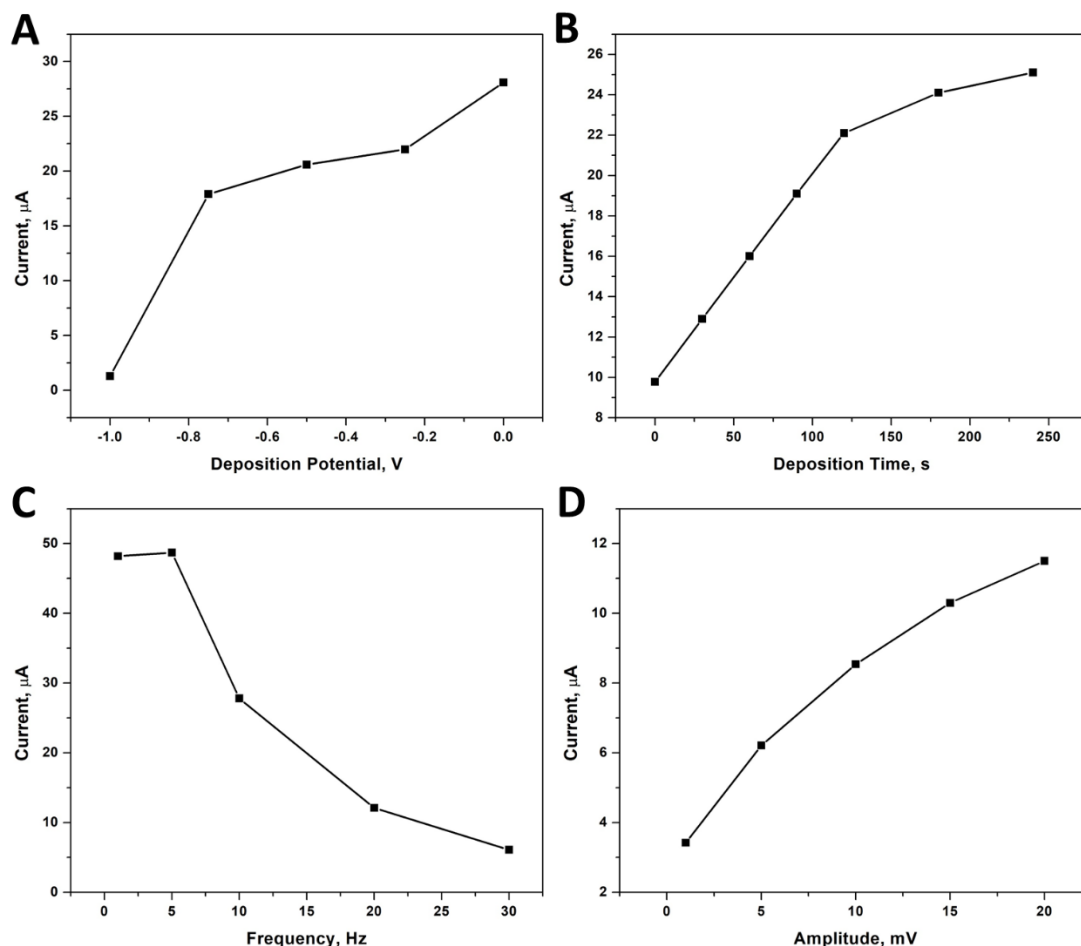
The influence of deposition potential on the stripping response of Ni^{2+} at the CB-DMG-Ag-PPPE was interrogated in the potential range between 0.0 and − 1.0 V (Figure 4A). Variations in the accumulation potential towards the established reduction potential of Ni^{2+} into Ni^0 (0 V to -0.8), showed a gradual decrease in recorded peak currents achieved from Ni^{2+} detection. Reduction of the metal cations present in solution converts metal cations to their neutral state lowering the overall concentrations of cations available for reduction in the cathodic sweep. A further move to − 1.0 V, sharply decreases the stripping peak currents. E_{acc} of -0.75 V was selected.

The effect of accumulation time on the Ni^{2+} stripping response was studied between 30 and 300 s. Figure 4B, shows a rapid increase in the Ni^{2+} peak current with increasing accumulation time between 0 and 180 s confirming the increase in adsorption of the Ni^{2+} on the CB-DMG-Ag-PPPE surface. Saturation of the electrode surface takes place at accumulation times greater than 180 s and results in a gradual stabilization in stripping peak current. A deposition time of 120 s was selected for all analyses.

Figure 4C shows the dependence of peak currents on the square wave frequency over the 0 to 30 Hz range. Maximum stripping voltammetric peak currents are achieved at low-frequency values indicating longer analysis times are required to improve the resolution and sensitivity of the paper-

based sensors. A steady decline in reduction peak currents between 5 and 30 Hz can be seen due to slower electron transfer processes through the DMG film. A frequency of 5 Hz was selected for the square-wave waveform.

The influence of amplitude on the stripping peak current of Ni^{2+} was studied between 0 and 20 mV and shown in Figure 4D. A steady increase in recorded stripping cathodic peak currents is observed.


Figure 4. *The influence of (a) accumulation potential (E_{acc}), (b) accumulation time (t_{acc}), (c) frequency, and (d) amplitude on Ni^{2+} peak currents at CB-DMG-Ag-PPPE. Sample composition: 300 µg L^{-1} Ni, 0.1 M NH$_3$/NH$_4$Cl buffer*

Quantitative analytical performance of the CB-DMG-Ag-PPPE

The favorable adsorptive stripping voltammetric performance of the CB-DMG-Ag-PPPE is demonstrated in Figure 5. Square-wave voltammograms and the calculated calibration plot recorded over 50 – 500 µg L^{-1} Ni^{2+} ions, at optimum conditions were performed in deaerated samples. The voltammograms were recorded between -0.9 and -1.3 V. Voltammograms show a single, broad, reversible cathodic stripping peak at ~ -1.18 V. Peak currents, credited to the reduction of Ni^{2+} cations from the [Ni(dmgH)$_2$] complex increase with increasing Ni^{2+} concentration. Positive peak potential shifts are indicated at higher metal ion concentrations. The amount of energy to convert the cation into its neutral state is therefore decreased. Figure 5, inset represents the noted calibration plot obtained from the corresponding voltammograms. A constant, linear increase in peak currents is observed between 0 and 350 µg L^{-1} [Ni(dmgH)$_2$] concentration. At Ni^{2+} concentrations greater than 350 µg L^{-1} saturation of the electrode surface with adsorbed metal-complex is experienced and a deviation from linear concentration is observed. A plateau is evident in the calibration curve. A dynamic linear range between 50 and 350 µg L^{-1} [Ni(dmgH)$_2$] concentration is established with a

correlation coefficient of 0.989 indicating good linearity in the recorded region at an accumulation time of 90 s. Table 1 is a summary of the analytical parameters obtained from the calibration curve.
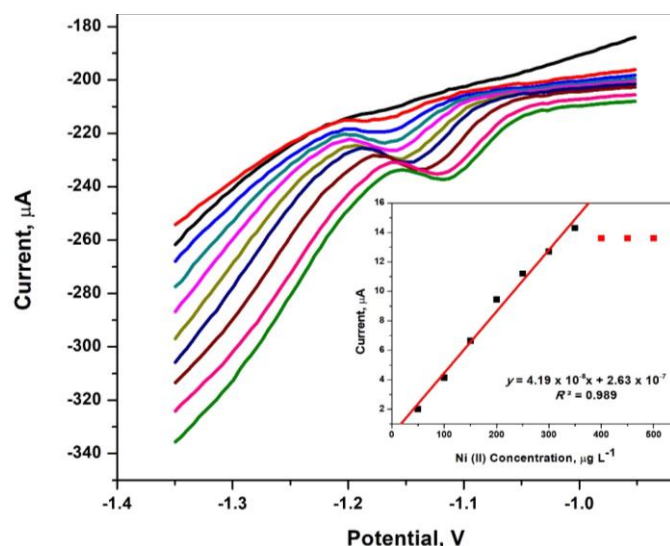

Figure 5. *SW-AdCSVs and corresponding calibration plot of individual analysis of Ni^{2+} obtained at CB-DMG-CC-Ag-PPE over a 50 – 500 µg L^{-1} range. Supporting electrolyte 0.1 NH_3/NH_4Cl Buffer, pH 9.4, deposition time 90 s, deposition potential - 0.7 V, rotation speed 1000 rpm, frequency 5 Hz, amplitude 0.01 V*

Table 1. *A summary of recorded analytical data for the CB-DMG-Ag-PPE over the 50 – 500 µg L^{-1} range*

Analytical Parameter	Analysis of Ni^{2+}
Sensitivity, µA L µg^{-1}	4.19 10^{-8}
Correlation coefficient	0.989
Detection limit ± SD, µg L^{-1}	48.01 ± 12.24
Limit of quantification ± SD, µg L^{-1}	144.03 ± 51.92

Analysis of tap water samples at the CB-DMG-Ag-PPPE

The carbon black, dimethylglyoxime paper-based sensor (CB-DMG-Ag-PPPE) was applied to the quantitative analysis of Ni^{2+} ions in tap water samples collected in our laboratory. Quantitative analysis of tap water was performed using a simple standard addition technique whereby successive additions of known sample concentrations were added to the sample under analysis to determine its concentration by extrapolation methods. The recorded voltammogram and standard addition calibration curve are shown in Figure 6.
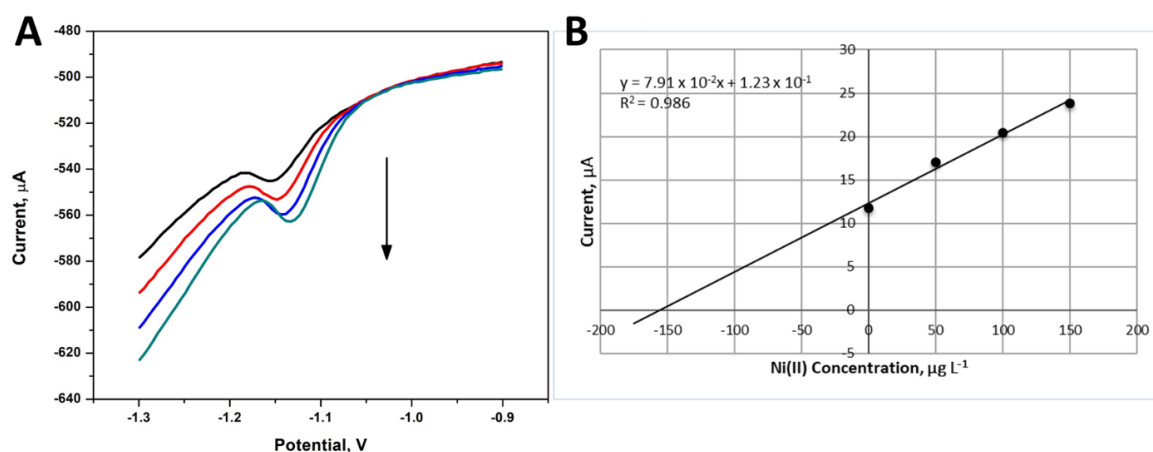
Figure 6. *Analysis of tap water (pH 9.4) spiked with 150 µg L^{-1} of the metal ion in the presence of 2 mM DMG and 10 mg L^{-1} Hg^{3+}. (a) SW-AdSV and (b) Calibration plot. Supporting electrolyte: 0.1 M NH_3/NH_4Cl buffer (pH 9.4). SWV instrumental parameters: E_{acc} = -0.7 V, t_{acc} = 120 s, f = 5 Hz and amplitude= 10 mV.*

A linear increase in peak current was observed with increasing "spike" concentration with good regression of 0.986. The extrapolated x-intercept was used to determine the concentration of Ni^{2+} cations in the unknown sample. A summary of the reported values for Ni^{2+} concentrations found in both the spiked and un-spiked samples are shown in Table 2, below. The samples were analysed before (original) and after (found) known concentrations (added) of Ni^{2+} were added. It was found that the Ni^{2+} ion concentration in tap water samples could not be determined as it was below the sensor limits of detection or was not present in the sample at all. Spiked samples allowed for accurate recovery studies in the new sample matrix when studied within the linear dynamic range found previously. Good recovery percentages were determined for Ni^{2+} detection in both test (buffer) and real tap water samples respectively with an error below 5 %.

Table 2. *Recorded recovery percentages in both test and water samples*

Ni²⁺ sample	c_{Ni} / µg L⁻¹			Recovery, %
	Original[1]	Added[2]	Found[3]	
Test sample	N/D	150	148.15	98.77 ± 7.41
Real water sample	N/D	150	155.49	103.66 ± 5.34

[1]Original – determined concentration of sample prior spiking with Ni^{2+}

[2]Added – concentration of Ni^{2+} spiked into the sample

[3]Found - determined concentration of sample after spiking with Ni^{2+}

A summary of the work performed at DMG modified electrodes is shown in Table 3, below. Limits of detection ranging from 0.5 to 48 µg L-1 were observed. Solid electrodes like glassy carbon and screen-printed electrodes provided the lowest limits of detection owing to the high conductivity of the solid substrates even at shorter analysis times. DMG modified carbon paste electrodes were also studied in select works. Previous work done in paper-based sensors by our research group is also shown. Paper-based electrochemical cells used in conjunction with SPCEs demonstrated slightly lower sensitivity due to the low sample volumes required for analysis, limiting the number of available ions for detection. As expected, the two works performed in paper-based sensors showed low sensitivity owing to the lack of electron transport in the cellulose matrix. The ERGO-AuNP-CC-Ag-PPPE in conjunction showed improved LOD because of the mercury film.

Table 3*. A summary of recent work performed at DMG modified electrodes*

Metal ions	Substrate	Technique	Accumulation time, s	Dynamic linear range, µg L⁻¹	Detection limit, µg L⁻¹	Ref.
Ni²⁺	DMG-CPE	DPAdSV	120	80 - 600	27	[2]
Ni²⁺	DMG-N-SPE	DPAdSV	120	60 - 500	30	[3]
Ni²⁺ with Co²⁺ & Zn²⁺	NGr-DMG-GCE	SWAdCSV	120	2 - 20	1.5	[4]
Ni²⁺	DMG-Hg-µPPEC	SWAdCSV	90	15 - 90	6.27	[5]
Ni²⁺	ERGO-AuNP-CC-Ag-PPPE	SWAdCSV	120	50 – 500	32.19	[6]
Ni²	DMG-SPCE	DPAdSV	120	1.7 - 150	0.5	[7]
Ni²	DMG-SPCE	SWAdSV	60	7.6 - 200	2.3	[8]
Ni²	DMG-CPE	DPAdSV	120	50 - 340	27	[9]
Ni²⁺	CB-DMG-Ag-PPPE	SWAdCSV	120	50 - 350	48.01	This work

Conclusion

This study describes a simple inkjet printing process to produce patterned three-electrode systems based on silver nanoparticles (AgNPs) deposited on commercial photographic paper without the need for stencils or templates. The use of nanoparticles in the electrode fabrication provided highly conductive and reproducible electrodes with uniform morphology. The fabricated

Ag-PPE showed good conductivity in the low ohm range. Further, modification of the Ag-PPE with a nafion carbon black dimethylglyoxime film demonstrated improved sensitivity and provided an accurate and simple quantitative analytical approach towards the detection of Ni^{2+} in drinking water samples by adsorptive cathodic stripping voltammetry (AdCSV). This was the first reported work on a carbon black dimethylglyoxime composite and the first use of carbon black as ink in paper-based sensing applications. The sensors provided good sensitivities and low limits of detection at short analysis times compared to similar studies on paper-based sensing and is only the second study showcasing paper-based devices in adsorptive stripping voltammetric applications. The work demonstrated a further approach to metal analysis at paper-based electrodes in the absence of commonly used metallic films and showcased comparable results to these. Accurate detection in the 98-103 % range was observed for test and real water samples. Further work is required to improve the low sensitivity of the paper-based substrates and investigate other nanostructured materials and fabrication processes for paper-based sensor development.

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