

1 **Modulated-laser source induction system for remote detection of**
2 **infrared emissions of high explosives using Laser-Induced Thermal**
3 **Emission (LITE)**

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17
18 **Abstract**

19 In a homeland security setting, the ability to detect explosives at a distance is a top security
20 priority. **Consequently, the development of remote, non-contact detection systems continues to**
21 **represent a path forward.** In this vein, a remote detection system for excitation of infrared
22 emissions using a CO₂ laser for generating Laser-Induced Thermal Emission (LITE) is a possible

23 solution. However, a LITE system using a CO₂ laser has certain limitations, such as the
24 requirement of careful alignment, interference by the CO₂ signal during detection, and the power
25 density loss due to the increase of the laser image at the sample plane with the detection distance.
26 In this study, a remote chopped-laser induction system for LITE detection using a CO₂ laser
27 source coupled to a focusing telescope was built to solve some of these limitations. Samples of
28 fixed surface concentration (500 μg/cm²) of 1,3,5-trinitroperhydro-1,3,5-triazine (RDX) were
29 used for the remote detection experiments at the distance ranging between 4 and 8 m. This
30 system was capable of thermally exciting and capturing the thermal emissions (TEs) at different
31 times in a cyclic manner by a Fourier Transform Infrared (FTIR) spectrometer coupled to a gold-
32 coated reflection optics telescope (FTIR-GT). This was done using a wheel blocking the capture
33 of TE by the FTIR-GT chopper while heating the sample with the CO₂ laser. As the wheel
34 moved, it blocked the CO₂ laser and allowed the spectroscopic system to capture the TEs of
35 RDX. Different periods (or frequencies) of wheel spin and FTIR-GT integration times were
36 evaluated to find dependence with observation distance of the maximum intensity detection,
37 minimum S/N ratio, CO₂ laser spot size increase, and the induced temperature increment (ΔT).

38

39 **Keywords**

40 Standoff detection; laser-induced thermal emission; highly energetic materials; mid-infrared
41 emission spectroscopy; carbon dioxide laser

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45 **1 Introduction**

46 Laser-Induced Thermal Emission (LITE)¹⁻⁷ detection of analytes is an important remote
47 detection technique for monitoring hazardous chemicals safely.⁸⁻²⁵ In a homeland security
48 setting, the detection of high explosives (HEs) enabled by LITE spectroscopy could play an
49 important role in **maintaining safe operations** while conducting real-time detection and fast
50 analysis.²⁶⁻³⁰ However, the limits of the detection (LOD) depend on the source to target distance,
51 on the laser power, on the divergence of the source, as well as on the collection efficiency and
52 detector sensitivity.³¹ Several studies have demonstrated that the use of a CO₂ laser as an
53 excitation source for LITE spectroscopy at long distances is highly suitable for stimulating
54 strong IR emissions of analytes of interest, such as trace amounts of HEs.^{1,12,32-34} This feasibility
55 of detection and identification of small quantities of HEs on surfaces is enabled by the fact that
56 CO₂ lasers operate in the mid-infrared (MIR), i.e., in the fingerprint region. HEs have strong
57 vibrational signatures in the MIR that can lead to generate resonant absorption by the analytes.
58 MIR radiation is strongly absorbed by the analyte, leading to relatively high local temperatures.
59 Thus, LITE spectroscopy is a technique that has many advantages: the sample itself is the source
60 of the analytical information,^{32,35} a strong resonant infrared photothermal approach,³¹ and the
61 facility for detecting other threat chemicals regardless of size, shape, and substrate type.

62 In a previous study developed by our research group, low surface concentrations of samples
63 of HEs were thermally excited to higher vibrational energy levels using a CO₂ laser, and the
64 radiant relaxation process was monitored using a telescope-coupled FTIR interferometer.^{1,36} The
65 emission spectra of HE: TNT, DNT, PETN, TATP, and RDX were successfully monitored at
66 distances from 4 to 64 m. During these experiments, several problems were identified. First, the
67 telescope should have been configured with a shutter to avoid blinding the detector due to the

68 back reflection of the laser beam. Second, the spectra were collected during the time interval in
69 which the sample was cooling after the laser was turned off. Third, the laser irradiance decreased
70 with the source-target distance because the laser spot (LS) size increased. The induced local
71 temperatures (ΔT) at each laser-target range (4, 8, 16, 32, and 64 m) were 72, 20, ~7, ~3, and > 1
72 °C. It was proposed that better results could be obtained by collimating the laser to increase the
73 ΔT and obtain higher emission levels. Fourth, careful alignment of the laser-target-detector angle
74 was necessary when the telescope and the excitation source are not part of the same device.
75 Therefore, the main objective of this study was to develop a new experimental set-up for remote
76 LITE detection of HEs and proposing a methodology that allows solving these four problems
77 that affect the experimental efficiency of LITE spectroscopy.

78

79 **2 Methodology**

80 *2.1 Sample preparation*

81 Only one HE was used in the experiments: 1,3,5-trinitroperhydro-1,3,5-triazine (RDX) which
82 was synthesized in the laboratory. Acetone (99.5%, GC grade, Thermo-Fisher Scientific,
83 Waltham, MA, USA) was used as the solvent. RDX samples were deposited on aluminum plates
84 substrates (Al: 2.54 cm \times 2.54 cm) at a nominal surface concentration of 500 $\mu\text{g}/\text{cm}^2$. A sample-
85 smearing technique was used to deposit the RDX samples.^{37,38} A stock solution of RDX was
86 prepared in acetone. A Teflon® stub (3 cm wide and 0.04 cm thick) was used to smear a fixed
87 volume (10 μL) of the solutions containing the analytes onto the metal surface. The smearing
88 was performed quickly, tilting the Teflon® sheet to the right or left in a single-pass operation. A

89 minimal adhesion of $500 \mu\text{g}/\text{cm}^2$ to the Teflon was assumed. The plates were first cleaned with
90 acetone followed by air-drying before the desired RDX surface loading was deposited.

91 2.2 Apparatus

92 The experimental setup used is illustrated in Fig. 1. It consisted of an intermittent excitation
93 system of the thermal emissions (TE) induced by a CO_2 laser as the excitation source. This
94 system was capable of thermally exciting and acquiring the TE at different times in a cyclic
95 manner using a Fourier transform infrared (FTIR) spectrometer coupled to a gold-coated optics
96 reflective telescope (GT). The chopping of the laser beam was accomplished using a wheel to
97 block the laser and the capture of the TE by the FTIR-GT while heating the sample with the CO_2
98 laser. As the wheel moved, it blocked the CO_2 laser and allowed the acquisition of the TEs by the
99 FTIR-GT. The sample temperature was measured with a thermocouple thermometer. On the
100 back of the aluminum plate that was used as sample substrate, a cylindrical cavity was drilled
101 close to the surface of the samples and the thermocouple head was fixed in place. This reduced
102 the temperature difference between the sample and thermocouple by less than $0.5 \text{ }^\circ\text{C}$.

103 2.2.1 Telescope

104 The catadioptric telescope used was an Argunov-Cassegrain design, equipped with gold-coated
105 mirrors. The primary and secondary mirrors were 15 cm and 5 cm in diameter, respectively. The
106 telescope was modified by allowing the secondary mirror to move and form an image at the focal
107 plane from infinite distance to less than 4 m. This was possible by placing the secondary mirror
108 in a movable optical mount (movement represented by a dotted arrow in Fig. 1) with a linear
109 displacement resolution of $1 \mu\text{m}$ to allow the decreased focal point of the reflective system.¹

110 2.2.2 Spectrometer

111 The infrared spectrometer used was an open-path-FT-IR interferometer, model EM27 (Bruker
112 Optics, Billerica, MA, USA). The optical bench consisted of a compact, enclosed, and desiccated
113 Michelson-type interferometer equipped with ZnSe windows, an internal blackbody (BB)
114 calibration source, a KBr beamsplitter, a very fast native focal ratio of $f/0.9$ with a field of view
115 (FOV) of 30 mrad (1.7°). The FOV was reduced to 10 mrad (0.6°) by the receiver telescope,
116 which had an infinite focal point magnification of $3\times$.¹

117 2.2.3 Optical chopper wheel (OCW) system

118 The optical chopper wheel (OCW) used consisted of an aluminum plate cut to a radius of 15.2
119 cm which had a two-and-a-half arc-shaped opening (slightly larger than the spectrometer
120 window) and an angle of 153.5° (Fig. 1(a)). The wheel contained a counterweight on the side of
121 the opening with its center of mass at the center of the wheel so that when turning, it would not
122 pivot.

123 *CO₂ laser*: The target surface was heated using a CO₂ laser with 36 W as maximum output
124 power (Firestar vi30, Synrad, Inc., Mukilteo, WA, USA), and an emission band from 926 to 980
125 cm^{-1} . The laser beam was aligned with the center of the spectrometer window in an imaginary
126 line called “the horizon.” The center of the telescope mirror was aligned with the center ray of
127 the laser beam. The laser radiation was reflected by two plane gold-coated mirrors to enter the
128 telescope. Then, the laser beam was reflected towards the secondary mirror of the telescope and
129 then to the primary mirror and then onto the sample (Fig. 1). This strategy of using the telescope
130 to focus the laser light ensures that the LS on the sample was small and the laser power was high
131 to ensure a strong heating effect. Thus, the telescope assumed two roles, focusing the laser and
132 collecting the emission. Moreover, the alignment with respect to the sample was minimal and
133 was only necessary to point at the sample plane. A 532 nm laser pointer was used to align the

134 path of the CO₂ laser to aim the sample. This part of the system contained a moving mirror that
135 allowed pointing to the sample. This mirror is removed to start the LITE detection.

136 *2.2.4 Time-step*

137 The OCW system divided the cycle into four-time steps (Fig. 1(b)). The first step was the
138 acquisition time. In this time, the OCW rotated 153.5° allowing the telescope to capture the
139 emissions and then passing through to the spectrometer and, at the same time, blocking the CO₂
140 laser. The second step was a transition time. In this period, the OCW opened to a 26.5° arc and
141 completely blocked the CO₂ laser letting the spectrometer window partially. In the third time
142 step, the OCW rotated by 153.5°, letting the laser light pass towards the telescope and blocking
143 the window of the spectrometer so that it did not detect the laser light when it was heating the
144 sample. Finally, in the fourth step consisted of a secondary transition time in which the OCW
145 turned by 26.5°, remaining to block the CO₂ laser while the spectrometer window was partially
146 closed.

147

148 **3 Results**

149 *3.1 Relative IR thermal emission of the OCW*

150 In the first set of experiments, samples of RDX of 500 µg/cm² nominal surface loading were
151 placed at a target distance 4 m from the cyclic induction system. The target samples were
152 focused by the telescope using the laser pointer. Next, the OCW was adjusted to a wheel
153 frequency (f_w) of 0.1 Hz, and the spectra were acquired with an integration time (t_i) of 2.0 s,
154 continuously. Fig. 2 (a) shows the profiles of the emission spectra. Two types of spectra were
155 acquired, emission spectra of the samples with the OCW switched off (Off), and sample

156 emission spectra of the samples with the OCW switched on (On). These are illustrated in Fig. 2
157 (a) by red and black traces, respectively. The ratios between these two types spectra were
158 calculated and called relative emission spectra (On/Off ratios or e_A) with respect to the aluminum
159 OCW at room temperature used as background.

160 In Fig. 2(a), the e_A spectrum of RDX (purple trace) reveals the characteristic spectral pattern
161 for RDX as can be observed in relation to an emissivity spectrum of RDX used reference to a
162 standard of RDX (green trace). Usually, the relative emission spectrum is obtained by dividing
163 the emission of the sample by the emission of the blackbody at the same temperature as the
164 sample. This situation generates an inverted signal spectrum since the blackbody emission is
165 always higher than the sample emission, but in this case, e_A spectra were obtained with respect to
166 the emission of the aluminum OWC. However, the chopper wheel material does not behave as a
167 blackbody, and its emission temperature is lower than the sample emission temperature. This has
168 the effect that the vibrational features in the spectrum are shown upward looking (as if it were an
169 absorption spectrum). This practical form of detection is convenient since, in the case of remote
170 detection in the field, the temperature of the sample is unknown. For that reason, this work has
171 been focused on making the detection considering the possibility of applications in real
172 scenarios.

173 The MIR signals for RDX that stand out include the band at 1234 cm^{-1} and the band at 1034
174 cm^{-1} for N-C-N stretch, the signals at 1016 cm^{-1} , and 943 cm^{-1} for the combination N-NO₂ stretch
175 and in-plane C-H bending, respectively. Other RDX signatures identified are 919 cm^{-1} out-of-
176 plane CH₂ bending, the ring breathing mode at 881 cm^{-1} , the band at 850 cm^{-1} for C-N-C stretch,
177 and the band at 779 cm^{-1} for N-C-N stretch.^{39,40} The band at 1321 cm^{-1} for N-NO₂ symmetric

178 stretch and the signal at 1593 cm^{-1} for the N-NO₂ asymmetric stretch were not observed because
179 water vapor signals overshadowed them (Fig 1(a)).

180 Fig. 2(b) shows the e_A spectra for 18 cycles of the OCW at $f_w = 0.1\text{ Hz}$. The inserted figure
181 shows the time step for this experiment, where the t_i was 2.0 s. The asterisks indicate the
182 spectrum used to obtain e_A . The increase of the characteristic emission RDX signals was
183 observed due to the increase in sample temperature.

184 *3.2 Signal monitoring of the OCW adjusted by frequency and integration time*

185 Samples of $500\text{ }\mu\text{g}/\text{cm}^2$ of RDX were placed at 4 m of distance from the cyclic thermal induction
186 system and focused onto the FOV of the telescope using the laser pointer. A background
187 spectrum of aluminum of OCW was initially acquired. Next, the f_w of the OCW system was
188 adjusted to 0.1, 0.2, 0.4, 1.0, and 2.0 Hz and acquiring spectra with a t_i of 2.0 s, continuously.
189 Then, the f_w was set at 1.0 Hz, and the t_i was varied between 0.5 to 2.2 s. In Fig. 3(a-d) the e_A for
190 the RDX signal at 1234 cm^{-1} at 100 s are shown, with a total of 50 spectra for different f_w values.
191 Three types of points are shown: blue diamond represent spectra in Off mode (when the wheel
192 obstructs the spectrometer window); red diamond represent spectra in On mode (when the
193 spectrometer window is clear), and green diamond are intermediate between On and Off
194 positions. In Fig. 3(d) when the spectra acquisition of the OCW was adjusted to $t_i = 2.0\text{ s}$ and f_w
195 2.0 Hz, the cyclic speed increase was evidenced. Therefore, the OCW would enable the
196 acquisition of four spectra in the On position and four spectra in the Off position. Consequently,
197 each green diamond represents the average of the e_A at 1234 cm^{-1} for RDX of all spectral
198 acquisitions in On and Off (On-Off) time. Fig 3 (a) shows only the e_A for the peak at 1234 cm^{-1}
199 in On and Off due to cyclic speed decrease.

200 Fig. 4(a) shows the e_A spectra #50 for different f_w in contrast with the standard spectra of
201 RDX (solid black line). In all e_A spectra, the characteristic signals of RDX were observed.
202 However, the best spectra e_A were observed at f_w values of 0.2 and 0.4 Hz due to the decrease of
203 the noise in the spectra. Although RDX signals in the e_A spectra for 0.1 Hz are very intense, they
204 also contain a higher noise compared to other spectra.

205 The effect of the integration time (t_i) was also evaluated. In this case, the cyclic system was
206 set at f_w 1.0 Hz and t_i was varied between 0.5 to 2.2 s. Fig. 4(b) shows the spectra with higher
207 intensity (max) and less intensity (min) and their spectral average (A_v), varying the t_i : 0.54 s,
208 1.08 s, and 2.16 s. In general, for all A_v spectra, the t_i values evaluated showed a variation in the
209 spectral pattern that represents the RDX spectrum was observed. However, the increase in t_i
210 allows that the OCW acquires enough sample information to get a good signal for both max and
211 min spectra.

212 *3.3 Dependence of CO₂ laser spot and ΔT with the distance*

213 The dependence of the LS diameter of the CO₂ laser and the induced temperature increment (ΔT)
214 by the LITE process on RDX samples was investigated. The LS diameter of the CO₂ laser was
215 determined to increase by 0.5 cm/meter without a focusing telescope (Galán-Freyte et al.¹). In
216 contrast, when the telescope was used to focus the LS the increase in diameter was only 0.1
217 cm/meter. LS diameter as a function of distance was measured using thermal paper (change in
218 color with temperature). For the measurements, thermal paper was exposed to the laser action for
219 5 s, leaving a stain of the size of the LS diameter. The measurements were repeated for 5 times
220 and an average was calculated. The uncertainty of the reported diameter values was ± 0.05 cm.
221 This reduced increment in LS diameter was reflected in the ΔT induced and the signal
222 dependence with source-target distance. The results are shown as insets of Fig. 5(a). Focusing

223 the LS with the telescope for LITE detection at 4 m induced a ΔT lower than LITE detection
224 without a telescope. This result is attributed to the fact that in the present optical arrangement,
225 the induced heat transfer is an intermittent process.

226 When the sensing distance increased to 8 m, the effect of the increase of the LS diameter per
227 meter is more evident. The ΔT for LITE detection without focusing the LS decreases
228 approximately 50 °C in comparison with focusing the LS in which the decrease in ΔT is
229 maintained almost to 10 °C, as shown on the right side inset of Fig. 5 (a). A fit was obtained
230 between the ΔT induced at the sample and the reciprocal of the LS diameter (shown on left side
231 inset of Fig. 5 (a)). The dependence obtained is as expected because when increasing the LS
232 diameter, the power per unit area decreases, generating a lower ΔT .

233 To evaluate the effect of the telescope as a focusing tool in remote detection using a LITE
234 intermittent induction system, a comparison was made with the report by Galán-Freyte *et al.*,¹
235 using the RDX samples deposited on Al substrates at the same nominal surface concentration
236 (500 $\mu\text{g}/\text{cm}^2$). The signal to noise ratio (S/N) for the RDX samples was measured and compared
237 between LITE systems with and without focusing the LS with the telescope. Fig. 5(b) shows the
238 averages of e_A for RDX spectra acquired with the cyclic induction LITE system at 4, 6, and 8 m
239 of source-target distances. The inset in Fig. 5(b) shows that at 4 m of source-target distance, a
240 higher S/N ratio was achieved by the non-focused LITE system. However, at 8 m of source-
241 target distance, the S/N ratio was higher for the cyclic induction LITE system focusing the LS
242 with the telescope system.

243

244 **4 Conclusion**

245 The proposed methodology showed an improved efficiency in the alignment of the LITE process
246 at the sample plane. The dependence of LITE with source-target distance was improved, solving
247 the excessive increase in the size spot of the CO₂ laser in the LITE process. The sample emission
248 rate was measured without the interference of the CO₂ laser light. The optimum conditions for
249 the cyclic induction LITE system were adjusted to 1 Hz for the frequency of the chopper wheel
250 and an integration time of 2.0 s. In this study, only LITE detection distances from 4 to 8 m were
251 evaluated. For future work, it is expected to investigate at a longer detection distances, with other
252 sources of excitation such as tunable MIR lasers (quantum cascade lasers), synchronizing the
253 detector with the chopper.

254 It is emphasized that it was not necessary to measure the temperature of a blackbody or to
255 have to turn off the laser because it saturated the detector. This is a definite advantage of cyclic
256 induction LITE systems making it more amenable real field applications.

257

258 **Disclosures**

259 The authors declare no conflicts of interest.

260

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268 **References**

269 1. N. J. Galán-Freyte et al., "Standoff Detection of Highly Energetic Materials Using Laser-
270 Induced Thermal Excitation of Infrared Emission," *Applied Spectroscopy* **69**(5), 535-544
271 (2015).

272 2. L. T. Lin, D. D. Archibald, and D. E. Honigs, "Preliminary Studies of Laser-Induced
273 Thermal Emission Spectroscopy of Condensed Phases," *Appl. Spectrosc.* **42**(3), 477-483
274 (1988).

275 3. W. M. Michael J. Wilhelm, Jonathan M. Smith, Hai Lung Dai, "The lowest quartet-state of
276 the ketyl (HCCO) radical: Collision-induced intersystem crossing and the ν_2 vibrational
277 mode," *Chemical Physics*. **422**(290-296 (2013)).

278 4. M. N. Michael J. Wilhelm, Laura Letendre, Hai Lung Dai, "Photodissociation of vinyl
279 cyanide at 193 nm: Nascent product distributions of the molecular elimination channels," *J.*
280 *Chem. Phys* **130**(4), (2009).

281 5. L. T. Letendre et al., "Time-Resolved FTIR Emission Spectroscopy of Transient Radicals," *J.*
282 *Chinese Chem* **52**(4), 677-686 (2005).

283 6. L. T. Letendre et al., "Interfacing a transient digitizer to a step-scan Fourier transform
284 spectrometer for nanosecond time resolved spectroscopy," *Rev. Sci. Instrum* **70**(1), 18-22
285 (1999).

286 7. V. Karpovych et al., "Laser-induced thermal emission of rough carbon surfaces," *Journal of*
287 *Laser Applications* **32**(1), 012010 (2020).

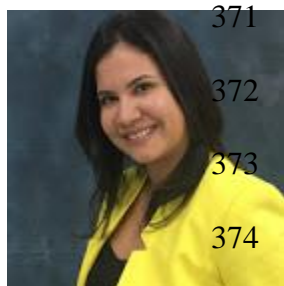
- 288 8. S. Wallin et al., "Laser-based standoff detection of explosives: a critical review," *Anal.*
289 *Bioanal. Chem* **395**(2), 259-274 (2009).
- 290 9. C. W. Van Neste, L. R. Senesac, and T. Thundat, "Standoff Spectroscopy of Surface
291 Adsorbed Chemicals," *Anal. Chem* **81**(5), 1952-1956 (2009).
- 292 10. L. Pacheco-Londoño et al., "Vibrational spectroscopy standoff detection of explosives," *Anal*
293 *Bioanal Chem* **395**(2), 323-335 (2009).
- 294 11. J. R. Castro-Suarez et al., "FT-IR Standoff Detection of Thermally Excited Emissions of
295 Trinitrotoluene (TNT) Deposited on Aluminum Substrates," *Appl Spectrosc* **67**(2), 181-186
296 (2013).
- 297 12. A. Mukherjee, S. Von der Porten, and C. K. N. Patel, "Standoff detection of explosive
298 substances at distances of up to 150 m," *Appl. Opt* **49**(11), 2072-2078 (2010).
- 299 13. J. C. Carter et al., "Standoff Detection of High Explosive Materials at 50 Meters in Ambient
300 Light Conditions Using a Small Raman Instrument," *Appl. Spectrosc.* **59**(6), 769-775 (2005).
- 301 14. J. L. Gottfried et al., "Standoff Detection of Chemical and Biological Threats Using Laser-
302 Induced Breakdown Spectroscopy," *Appl. Spectrosc.* **62**(4), 353-363 (2008).
- 303 15. A. K. Misra et al., "Single-Pulse Standoff Raman Detection of Chemicals from 120 m
304 Distance During Daytime," *Appl. Spectrosc.* **66**(11), 1279-1285 (2012).
- 305 16. J. E. Parmeter, "The challenge of standoff explosives detection," *Security Technology, 2004.*
306 *38th Annual 2004 International Carnahan Conference on* 355-358 (2004).
- 307 17. B. E. Bernacki, and M. C. Phillips, "Standoff hyperspectral imaging of explosives residues
308 using broadly tunable external cavity quantum cascade laser illumination," *Proc. SPIE*
309 **7665**(76650I-76650I-76610 (2010).

- 310 18. W. Ortiz-Rivera et al., "Vibrational spectroscopy standoff detection of threat chemicals,"
311 *Proc. SPIE* **8031**(803129-803129-803110 (2011).
- 312 19. A. Pettersson et al., "Explosives standoff detection using Raman spectroscopy: from bulk
313 towards trace detection," *Proc. SPIE* **7664**(76641K-76641K-76612 (2010).
- 314 20. A. Pettersson et al., "Near Real-Time Standoff Detection of Explosives in a Realistic
315 Outdoor Environment at 55 m Distance," *Propellants Explos. Pyrotech* **34**(4), 297-306
316 (2009).
- 317 21. R. D. W. Alan R. Ford, Darius M. Vunck, Jeremy B. Rose, Thomas B. Blank, Ken R. Pohl,
318 Edwin L. Dottery, Troy A. McVay, Mikella E. Hankus, Ellen L. Holthoff, Paul M.
319 Pellegrino, Steve D. Christesen, Augustus W. Fountain III "Explosives Sensing Using
320 Multiple Optical Techniques in a Standoff Regime with a Common Platform,"
321 *Spectroscopyonline* 6-11 (April 2011).
- 322 22. N. J. Galán-Freyte et al., "Artificial Intelligence Assisted Mid-Infrared Laser Spectroscopy In
323 Situ Detection of Petroleum in Soils," *Applied Sciences* **10**(4), 1319 (2020).
- 324 23. G. L. McEneff et al., "Sorbent Film-Coated Passive Samplers for Explosives Vapour
325 Detection Part B: Deployment in Semi-Operational Environments and Alternative
326 Applications," *Scientific Reports* **8**(1), 5816 (2018).
- 327 24. W. Zhang et al., "Recent Developments in Spectroscopic Techniques for the Detection of
328 Explosives," *Materials* **11**(8), 1364 (2018).
- 329 25. F. Jin et al., *Chemical and explosive detection with long-wave infrared laser induced*
330 *breakdown spectroscopy*, SPIE (2016).
- 331 26. R. J. Pell et al., "Quantitative infrared emission spectroscopy using multivariate calibration,"
332 *Anal. Chem* **60**(24), 2824-2827 (1988).

- 333 27. M. Friedrich, and D. R. T. Zahn, "Emission Spectroscopy: An Excellent Tool for the Infrared
334 Characterization of Textile Fibers," *Appl. Spectrosc.* **52**(12), 1530-1535 (1998).
- 335 28. M. J. Zuerlein et al., "Modeling thermal emission in dental enamel induced by 9–11 μm laser
336 light," *Appl. Surf. Sci* **127–129**(0), 863-868 (1998).
- 337 29. R. W. Jones et al., "Chemical Analysis of Wood Chips in Motion Using Thermal-Emission
338 Mid-Infrared Spectroscopy with Projection to Latent Structures Regression," *Anal. Chem*
339 **74**(2), 453-457 (2001).
- 340 30. T. M. Niemczyk, S. Zhang, and D. M. Haaland, "Monitoring Dielectric Thin-Film
341 Production on Product Wafers Using Infrared Emission Spectroscopy," *Appl. Spectrosc.*
342 **55**(8), 1053-1059 (2001).
- 343 31. R. Furstenberg et al., "Stand-off detection of trace explosives via resonant infrared
344 photothermal imaging," *Appl. Phys. Lett* **93**(22), - (2008).
- 345 32. N. Y. Galán-Freyte et al., "Standoff laser-induced thermal emission of explosives," *Proc.*
346 *SPIE* **8705**(870508-870508-870510 (2013).
- 347 33. A. Figueroa-Navedo et al., "Improved detection of highly energetic materials traces on
348 surfaces by standoff laser-induced thermal emission incorporating neural networks," *Proc*
349 *SPIE* **8705**(87050D-87050D-87011 (2013).
- 350 34. A. M. Figueroa-Navedo et al., "Chemometrics-enhanced laser-induced thermal emission
351 detection of PETN and other explosives on various substrates," *Journal of Chemometrics*
352 **29**(6), 329-337 (2015).
- 353 35. F. B. Gonzaga, and C. Pasquini, "Near-Infrared Emission Spectrometry Based on an
354 Acousto-Optical Tunable Filter," *Anal. Chem* **77**(4), 1046-1054 (2005).

- 355 36. A. Tsuge, Y. Uwamino, and T. Ishizuka, "Applications of Laser-Induced Thermal Emission
356 Spectroscopy to Various Samples," *Appl. Spectrosc.* **43**(7), 1145-1149 (1989).
- 357 37. O. Primera-Pedrozo et al., "High Explosives Mixtures Detection Using Fiber Optics
358 Coupled: Grazing Angle Probe/Fourier Transform Reflection Absorption Infrared
359 Spectroscopy," *Sens Imaging* **9**(3-4), 27-40 (2008).
- 360 38. M. Wrable-Rose et al., "Preparation of TNT, RDX and Ammonium Nitrate Standards on
361 Gold-on-Silicon Surfaces by Thermal Inkjet Technology," *Sens Imaging* **11**(4), 147-169
362 (2010).
- 363 39. R. Infante-Castillo, L. C. Pacheco-Londoño, and S. P. Hernández-Rivera, "Monitoring the
364 $\alpha \rightarrow \beta$ solid-solid phase transition of RDX with Raman spectroscopy: A theoretical and
365 experimental study," *J. Mol. Struct* **970**(1-3), 51-58 (2010).
- 366 40. R. Infante-Castillo, L. Pacheco-Londoño, and S. P. Hernández-Rivera, "Vibrational spectra
367 and structure of RDX and its ¹³C- and ¹⁵N-labeled derivatives: A theoretical and
368 experimental study," *Spectrochim. Acta, Part A: Molecular and Biomolecular Spectroscopy*
369 **76**(2), 137-141 (2010).

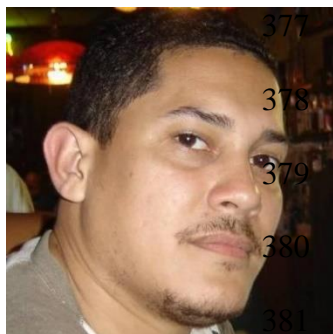
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430 **FIGURE CAPTIONS**

431 **Fig. 1** Experimental setup of the standoff laser-induced thermal detection (LITE) of HEMs. (a)
432 Optical chopper wheel (OCW) system; (b) Schematic representation of experiment timing for the
433 acquisition time and LITE.

434

435 **Fig. 2** (a) Thermal emission for $500 \mu\text{g}/\text{cm}^2$ at 4 m of RDX with $f_w = 0.1 \text{ Hz}$, $d = 4 \text{ m}$, and $t_i = 2$
436 s; (b) e_A spectra for 18 cycles at 1 Hz and 4 m.

437

438 **Fig. 3** Signal monitoring at 1234 cm^{-1} for (a) $f_w = 0.1 \text{ Hz}$, $t_i = 2.0 \text{ s}$, and $d = 4 \text{ m}$; (b) for $f_w = 0.2$
439 Hz, $t_i = 2 \text{ s}$, and $d = 4 \text{ m}$; (c) for $f_w = 0.4 \text{ Hz}$, $t_i = 2.0 \text{ s}$, and $d = 4 \text{ m}$; (d) for $f_w = 2.0 \text{ Hz}$, $t_i = 2.0 \text{ s}$,
440 and $d = 4 \text{ m}$.

441

442 **Fig. 4** (a) e_A spectra for various f_w values for $t_i = 2.0 \text{ s}$ and $d = 4 \text{ m}$; (b) average e_A spectra for
443 various t_i for a total acquisition time of 120 s, $f_w = 1.0 \text{ Hz}$ and $d = 4 \text{ m}$. The min and max spectra
444 of the experiment are shown.

445

446 **Fig. 5** (a) Dependence of the LS diameter for various source-target distances, induced with and
447 without the telescope. Insets: dependence of the temperature increase (ΔT) for various source-
448 target distances (left); reciprocal of LS diameter (right); (b) e_A spectra for various source-target
449 distances for $500 \mu\text{g}/\text{cm}^2$. Inset: dependence of S/N ratio divided by the nominal surface
450 concentration (C_s) for various source-target distances.

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