

***Ab initio* Molecular Orbital Studies of the Vibrational Spectra of some van der Waals Complexes. Part 4. Complexes of Sulphur Dioxide with Carbon Dioxide, Carbonyl Sulphide, Carbon Disulphide and Nitrous Oxide**

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ABSTRACT

The binary complexes formed between sulphur dioxide, as electron donor, and the series carbon dioxide, carbonyl sulphide and carbon disulphide, as electron acceptors, have been studied by means of *ab initio* molecular orbital theory. The optimized structures, the interaction energies and the vibrational spectra have been determined, and the effect of successive substitution of sulphur for oxygen atoms in the electron acceptor molecules has been established. Nitrous oxide, which is isoelectronic with carbon dioxide, has also been included among the electron acceptors, but the properties of the complex formed between sulphur dioxide and nitrous oxide are substantially different from those of the other three complexes.

KEYWORDS

Ab initio calculations, molecular complexes, sulphur dioxide, carbon dioxide, carbonyl sulphide, carbon disulphide, nitrous oxide, molecular structures, interaction energies, vibrational spectra.

1. Introduction

We have reported a series of *ab initio* studies of a variety of molecular complexes containing the family of Lewis acids carbon dioxide, carbonyl sulphide and carbon disulphide, with the aim of determining the effect of successively substituting sulphur for oxygen atoms in the electron acceptor molecules. The range of electron donors examined in these studies includes dinitrogen,¹ carbon monoxide,² ozone,³ water and hydrogen sulphide.^{4,5} The corresponding complexes containing sulphur dioxide as the electron donor have been studied experimentally, by Fourier transform microwave spectroscopy.^{6–8} Nitrous oxide, being isoelectronic with carbon dioxide, might be expected to form a complex with sulphur dioxide whose properties were similar to those of the three main subjects of this work. However, our studies have shown that the N₂,N₂O complex exhibited quite different properties from those of the N₂,CO₂, N₂,OCS and N₂,CS₂ analogues, and this was interpreted as being due to the fundamentally different electronic structures of N₂O and the other three acceptor monomer species.¹ The SO₂,N₂O complex, investigated by Fourier transform microwave spectroscopy,⁹ was found to have a different structure from those between SO₂ and CO₂,⁶ OCS⁷ and CS₂,⁸ determined by the same group of workers under comparable conditions. To the best of our knowledge, among the complexes of SO₂ considered here, only SO₂,N₂O has been subjected to an *ab initio* treatment.¹⁰ In this paper we describe the results of our computational studies of the complexes SO₂,CO₂, SO₂,OCS and SO₂,CS₂, and attempt to rationalize the similarities and differences among their properties, and then go on to present the corresponding results obtained for the anomalous SO₂,N₂O adduct and discuss the reasons for the different nature of this species.

2. Computational Methods

The calculations were performed using the Gaussian-98 computer program.¹¹ The level of treatment employed was second order Møller-Plesset perturbation theory (MP2),¹² and the basis set used was the 6-311+G(d) basis set of Pople¹³ and Schleyer.¹⁴ The structures were optimized, initially using the experimentally-determined symmetries^{6–9} and then, when those structures were found not to be genuine minima, by successively relaxing the geometrical constraints. The VERYTIGHT convergence criterion¹¹ was used, in order to ensure as close an approach to the real minimum energy structures as possible. The interaction energies were computed from the energies of the complexes and of the relaxed monomers, in their structures as found in the complexes, and were corrected for basis set superposition error (BSSE),¹⁵ using the full counterpoise technique of Boys and Bernardi,¹⁶ and for zero-point energy differences.

3. Results and Discussion

3.1. Molecular Structures

The geometries of the SO₂,CO₂ and SO₂,CS₂ complexes were first optimized under C_{2v} symmetry, in accordance with the symmetry of the former species determined experimentally by Kuczkowski *et al.*⁶ However, these structures were found to be first-order (SO₂,CO₂) and second-order (SO₂,CS₂) saddle points. The symmetries were then reduced to C_s, and the optimizations repeated. Although an energy saving was achieved in each case, SO₂,CO₂ was still found to be a first-order saddle point, while SO₂,CS₂ changed to a first-order transition state. Finally, both structures were re-optimized with all symmetry constraints removed, and both species converged with no imaginary modes. The SO₂,OCS and SO₂,N₂O complexes were optimized, starting with C_s symmetry,^{7,9} with the SO₂ molecule straddling

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Table 1 Energies (in Hartree) and Hessian indices (in brackets) of the complexes.

Complex	C_{2v}	C_s	C_1
SO ₂ .CO ₂	-735.99842846 (1)	-735.99860971 (1)	-735.99891746 (0)
SO ₂ .OCS (structure 1)	–	-1058.56470338 (1)	-1058.56488105 (0)
SO ₂ .OCS (structure 2)	–	– ^a	-1058.56421514 (0)
SO ₂ .CS ₂	-1381.13099425 (2)	-1381.13200018 (1)	-1381.13224863 (0)
SO ₂ .N ₂ O (structure 1)	–	-732.08245668 (1)	-732.08266382 (0)
SO ₂ .N ₂ O (structure 2)	–	-732.08254352 (0)	-732.08264767 (0)

^a Relaxed to structure 1.

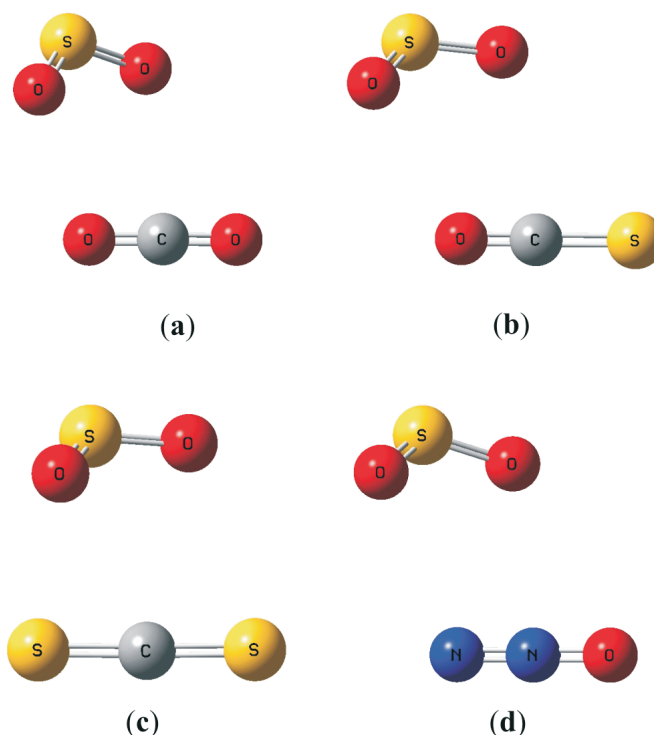
the plane containing the S (SO₂), O, C and S (OCS), or S (SO₂), N, N and O (N₂O) atoms. Two isomers were considered, in which the S atom of SO₂ approached the O (structure 1) and the S atom of OCS (structure 2), or the terminal N (structure 1) and the O atom of N₂O (structure 2). Structure 1 of SO₂.OCS was the preferred experimental structure,⁷ while in the case of SO₂.N₂O the observed structure lacked a symmetry plane.⁹ In our calculations, structure 1 of SO₂.OCS optimized with one imaginary vibrational mode, while structure 2 relaxed to structure 1. For SO₂.N₂O, structure 1 was found to be a first-order saddle point, while structure 2, which was slightly lower in energy, was a true minimum. Finally, optimization under C₁ symmetry resulted in all modes being real, with an energy saving in each case. These results are collected in Table 1.

The barrier heights to interconversion among the various isomers are given in Table 2. These barriers imply that in most cases the wavenumber of the interconversion mode is of the order of 50 cm⁻¹, except for the C_{2v} → C_s mode of SO₂.CS₂, where it is about 220 cm⁻¹. Structure 1 of SO₂.OCS, which, for want of a better description, we shall call the S...O bonded species, and structure 1 of SO₂.N₂O (the S...N bonded species) are clearly the most stable structures of these two complexes, and further discussion will be confined to these isomers. The preferred minimum energy structures of all four complexes are illustrated in Fig. 1, and the Cartesian coordinates are available as Supplementary Material.

While Kuczkowski *et al.* identified the symmetry of SO₂.CO₂ as C_{2v}, they conceded that, due to the large uncertainty in their determination of the A rotational constant, they could not confirm this structure unequivocally.⁶ We interpret our finding that the true symmetry is C₁ by proposing that the experimentally observed structure is a vibrationally-averaged one; the C_{2v} structure is a saddle point connecting two equivalent C_s structures in which the C₂ axis of the SO₂ molecule is tilted slightly relative to the S...C axis. The two C_s structures are interconverted by a SO₂ wagging vibration, coupled with an in-plane OCO libration. The form of this imaginary vibrational mode is illustrated in Fig. 2. Moreover, the C_s structure is also a saddle point, linking two equivalent C₁ forms which are interconverted

Table 2 Barrier heights to interconversion among the various isomers of the complexes.

Complex	Interconversion	Barrier height/kJ mol ⁻¹
SO ₂ .CO ₂	C _{2v} → C _s	0.476
	C _s → C ₁	0.808
SO ₂ .OCS	C _s → C ₁	0.466
SO ₂ .CS ₂	C _{2v} → C _s	2.641
	C _s → C ₁	0.652
SO ₂ .N ₂ O (structure 1)	C _s → C ₁	0.544
SO ₂ .N ₂ O (structure 2)	C _s → C ₁	0.273

**Figure 1** The optimized structures of the C₁ complexes of SO₂ with (a) CO₂, (b) OCS, (c) CS₂ and (d) N₂O.

through a SO₂ rocking mode, coupled with an out-of-plane OCO libration, as shown in Fig. 3.

A potential surface scan of the OS...CO dihedral angle confirms the presence of an energy barrier at the C_{2v} structure (see

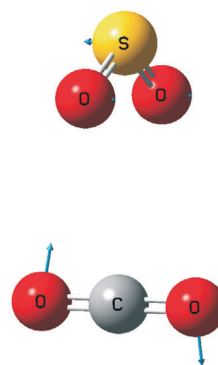
**Figure 2.** The form of the imaginary mode of the C_{2v} structure of the SO₂.CO₂ complex.

Fig. 4). A scan of the OSO...C dihedral angle revealed that a similar barrier exists, for interconversion of the two C_1 species, as illustrated in Fig. 5. For the $\text{SO}_2\cdot\text{CS}_2$ C_{2v} complex the two imaginary modes are the counterparts of those represented in Figs 2 and 3.

In the case of $\text{SO}_2\cdot\text{CS}_2$, Kuczkowski *et al.* determined a C_s structure, with two alternative orientations fitting the experimental spectrum equally well.⁸ Our interpretation of the true structure involving interconversion of two C_1 species through a small C_s barrier holds here too. The experimental structure of $\text{SO}_2\cdot\text{OCS}$ is C_s , with the C_2 axis of SO_2 almost parallel to the OCS molecule, and with their dipole moments opposed.⁷ For $\text{SO}_2\cdot\text{N}_2\text{O}$ the experimental symmetry is C_1 , as evidenced by the splitting of some of the rotational transitions into doublets, arising from a tunnelling motion.⁹ As a result the complex lacks a symmetry plane, in agreement with our calculated structure. The authors

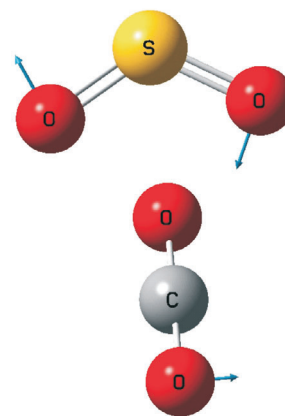


Figure 3 The imaginary mode of the C_s structure of the $\text{SO}_2\cdot\text{CO}_2$ complex.

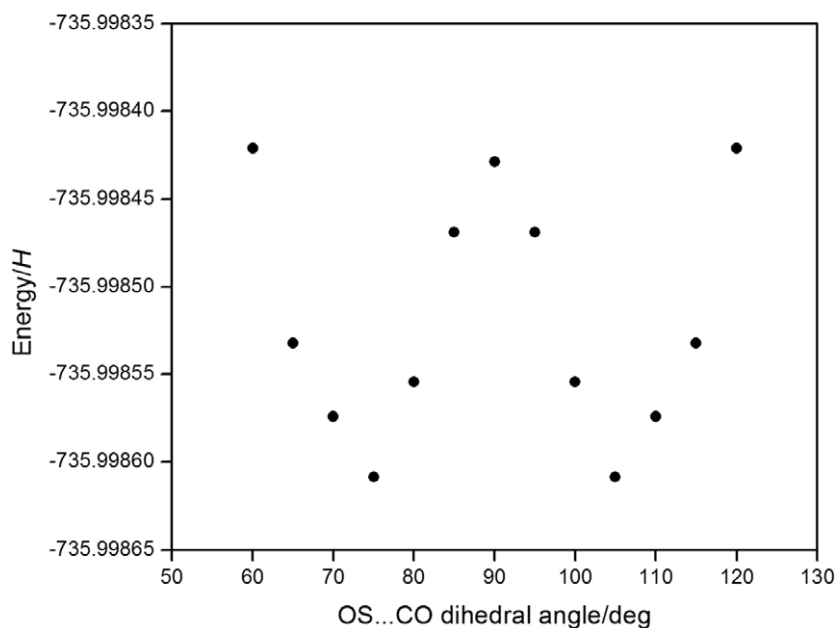


Figure 4 Potential surface scan of the OS...CO dihedral angle of the $\text{SO}_2\cdot\text{CO}_2$ complex of C_s symmetry.

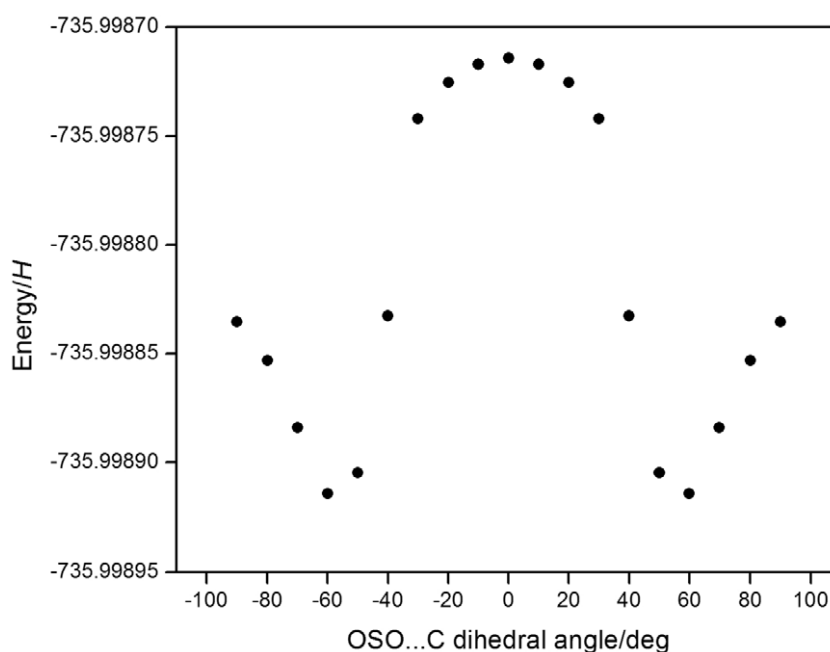


Figure 5 Potential surface scan of the OSO...C dihedral angle of the $\text{SO}_2\cdot\text{CO}_2$ complex of C_1 symmetry.

Table 3 Natural bond orbital charges of the interacting monomers.

Complex	Charge/e	
	SO ₂ fragment	Partner fragment
SO ₂ .CO ₂	0.0025	-0.0025
SO ₂ .OCS	0.0020	-0.0020
SO ₂ .CS ₂	-0.0071	0.0071
SO ₂ .N ₂ O	0.0007	-0.0007

Table 4 Primary and secondary orbital interactions of the complexes.

Complex	Primary interaction	Secondary interaction
SO ₂ .CO ₂	SO ₂ n(O) → CO ₂ σ*(CO)	CO ₂ π(nb) → SO ₂ π*(SO)
SO ₂ .OCS	SO ₂ n(O) → OCS σ*(CS)	OCS π(CS) → SO ₂ π*(SO)
SO ₂ .CS ₂	CS ₂ π(nb) → SO ₂ π*(SO)	SO ₂ n(O) → CS ₂ π*(CS)
SO ₂ .N ₂ O	SO ₂ n(O) → N ₂ O σ*(NO)	N ₂ O π(NN) → SO ₂ π*(SO)

failed to observe any such splittings in the other three complexes; this is clearly associated with the lower symmetry of SO₂.N₂O. In SO₂.OCS and SO₂.N₂O also, we envisage the tunnelling to involve mainly a SO₂ rocking motion, interconverting two equivalent C₁ structures. Valdes and Sordo have also computed the structure of the SO₂.N₂O complex.¹⁰ They examined six separate models and observed that all six were close in energy, with a very flat potential energy surface, and also invoked dynamic structure averaging to explain their results.

The interactions responsible for the stabilization of the four complexes may be pictured as an interplay between a primary donation of charge from an oxygen lone pair orbital associated with the SO₂ sub-molecule to a σ* orbital located mainly on the central atom of the electron acceptor, and a secondary back-donation from a π type orbital associated with the closer terminal atom of the acceptor to the π* orbital of the SO₂ molecule, which projects essentially normally to the SO₂ molecular plane. However, while a natural bond order (NBO)¹⁷ charge analysis indicates that SO₂ is a net electron donor to CO₂ and OCS, the reverse is true for SO₂.CS₂, and the charge transfer in the case of SO₂.N₂O is negligible (see Table 3).

The reversal of behaviour between CO₂ and OCS as electron acceptors on the one hand and CS₂ as a net donor on the other may be traced to the change of sign of the molecular quadrupole moment on progressing from CO₂ through OCS to CS₂ (-14.3 × 10⁻⁴⁰ C m² for CO₂, -3.0 × 10⁻⁴⁰ C m² for OCS and 11.5 × 10⁻⁴⁰ C m² for CS₂).^{18,19} The virtual lack of charge transfer in the case of

Table 5 Intramolecular structural parameters of the complexes, and their changes relative to those of the monomers.

Complex	Parameter	Value	Difference
SO ₂ .CO ₂	r(CO)/pm	117.11	0.11
		116.85	-0.15
SO ₂ .OCS	r(SO)/pm	146.83	0.09
		146.94	-0.03
	∠OSO/deg	118.88	-0.38
	r(CO)/pm	117.23	0.18
SO ₂ .CS ₂	r(CS)/pm	155.96	-0.32
	r(SO)/pm	146.92	0.06
		146.85	-0.01
	∠OSO/deg	118.96	-0.30
SO ₂ .N ₂ O	r(CS)/pm	156.54	0.38
		155.83	-0.34
	r(SO)/pm	147.06	0.20
		146.89	0.03
SO ₂ .N ₂ O	∠OSO/deg	118.89	-0.36
	r(NN)/pm	116.35	0.06
	r(NO)/pm	117.89	-0.22
	r(SO)/pm	146.99	0.13
	146.84	-0.02	
	∠OSO/deg	118.88	-0.38

SO₂.N₂O can be attributed to the unfavourable alignment of the axes of the terminal nitrogen lone pair of N₂O (along the molecular axis) and of the π* orbital of SO₂. The orbital interactions are summarized in Table 4.

The changes to the intramolecular structural parameters on complexation are reported in Table 5. While the changes are minimal, in general the SO bond lying in the pseudoplane containing the linear acceptor molecule and the S atom of SO₂ lengthens slightly while the SO bond which projects out of the pseudoplane shortens. The OSO angle consistently reduces by about a third of a degree. The bond of the electron acceptor closer to the S atom of SO₂ lengthens in each case and the bond remote from the interaction shortens, indicating a weakening of the intramolecular bond associated with the interaction and a strengthening of the more remote intramolecular bond.

Table 6 lists the computed intermolecular structural properties of the complexes. The most characteristic of these properties is the separation of the O atom of the SO₂ molecule lying in the pseudoplane and the central atom of the electron acceptor. In the series CO₂ – OCS – CS₂ this distance increases regularly, indicating a successively weaker interaction along the series. For SO₂.N₂O the N...O distance is similar to the C...O separation in SO₂.CO₂, suggesting a similarity in the properties of these two

Table 6 Intermolecular structural parameters of the complexes.

Complex	Parameter	Value	Complex	Parameter	Value
SO ₂ .CO ₂	R(S...O)/pm	355.21	SO ₂ .CS ₂	R(S...S)/pm	363.04
	R(C...O ⁱ)/pm ^a	290.42		R(C...O ⁱ)/pm ^a	330.55
	∠OC...O ⁱ /deg ^a	83.45		∠SC...O ⁱ /deg ^a	101.64
	∠O ⁱ S...O/deg ^a	56.60		∠O ⁱ S...S/deg ^a	90.66
	∠O ^o S...O/deg ^b	86.75		∠O ^o S...S/deg ^b	94.45
SO ₂ .OCS	R(S...O)/pm	333.04	SO ₂ .N ₂ O	R(S...N)/pm	350.15
	R(C...O ⁱ)/pm ^a	305.93		R(N...O ⁱ)/pm ^a	293.54
	∠OC...O ⁱ /deg ^a	79.78		∠NN...O ⁱ /deg ^a	85.31
	∠O ⁱ S...O/deg ^a	67.22		∠O ⁱ S...N/deg ^a	60.88
	∠O ^o S...O/deg ^b	85.62		∠O ^o S...N/deg ^b	86.91

^a Oⁱ – O atom of SO₂ lying in the pseudoplane.

^b O^o – O atom of SO₂ out of the pseudoplane.

Table 7 Comparison of computed and experimental intermolecular structures of the complexes.

Complex	Parameter	Computed	Experimental ^a
SO ₂ CO ₂	R _{cm} /pm ^b	362.23	329
	∠O1C...X1/deg ^c	66.96	
	∠C...X1...S/deg ^c	129.83	
SO ₂ OCS	R _{cm} /pm ^b	384.56	374.71
	∠S1...X2...X3/deg ^d	115.81	123.8
	∠X2...X3...S2/deg ^d	152.79	142
SO ₂ CS ₂	R _{cm} /pm ^b	354.26	342.87
	∠S3C...X4/deg ^e	96.22	92.7
	∠C...X4...S4/deg ^e	96.65	102.5
SO ₂ N ₂ O	R _{cm} /pm ^b	358.39	333.06
	∠N1...X5...X6/deg ^f	76.25	111.1
	∠X5...X6...S/deg ^f	118.16	156.9

^a See refs. 6–9.^b R_{cm} = centre of mass separation.^c O1 = O atom of CO₂ closer to S atom, X1 = centre of mass of SO₂.^d S1 = S atom of OCS, S2 = S atom of SO₂, X2 = centre of mass of OCS, X3 = centre of mass of SO₂.^e S3 = S atom of CS₂ farther from S atom of SO₂, X4 = centre of mass of SO₂, S4 = S atom of SO₂.^f N1 = terminal N atom of N₂O, X5 = centre of mass of N₂O, X6 = centre of mass of SO₂.**Table 8** Interaction energies of the complexes, corrected for basis set superposition error (BSSE) and for zero-point energy differences.

Complex	Interaction energy/kJ mol ⁻¹	
	Corrected for BSSE only	Corrected for BSSE and ΔE ₀
SO ₂ CO ₂	-4.119	-2.375
SO ₂ OCS	-3.207	-1.643
SO ₂ CS ₂	-2.650	-1.142
SO ₂ N ₂ O	-4.799	-3.025

complexes, but implying a significant difference in the natures of the CO₂ – OCS – CS₂ set on one hand and the N₂O species on the other. The data presented in Table 6 are not directly comparable with the reported experimental intermolecular geometrical properties,^{6–9} and our distances and angles have therefore been used to calculate the centre-of-mass separations and some of the angles which may be more sensibly compared with the experimental counterparts, as shown in Table 7. Because of the different symmetries of our computed structures and those (with the exception of SO₂N₂O) observed experimentally, the results shown in Table 7 understandably do not exhibit close agreement. Generally, our centre-of-mass separations overestimate the experimental values by between 2 and 11 %, but the intermolecular angles show no such consistency.

3.2. Interaction Energies

Table 8 presents the interaction energies of all four complexes, corrected for BSSE and also for zero-point energy differences. The counterpoise corrections were applied to the energies of the optimized structure complexes only, and no attempt was made to include such corrections in the optimization process. Both sets of energies indicate that all the complexes are bound, although only very weakly. The interaction energies decrease in the order N₂O > CO₂ > OCS > CS₂. In common with our observations regarding other complexes containing CO₂, OCS and CS₂,^{1–5} our interaction energies are found to be directly proportional to the ionization energy,^{20a} the absolute electronegativity²¹ and the

Table 9 Intramolecular vibrational wavenumbers of the complexes and their shifts relative to those of the monomers. All modes are of *a* symmetry.

Complex	Mode	Wavenumber /cm ⁻¹	Wavenumber shift/cm ⁻¹	Approximate description
SO ₂ CO ₂	ν ₁	2434.9	2.1	ν _a (CO ₂)
	ν ₂	1337.4	1.9	ν _s (CO ₂)
	ν ₃	1285.3	-0.3	ν _a (SO ₂)
	ν ₄	1075.3	2.5	ν _s (SO ₂)
	ν ₅	656.2	1.0	γ(CO ₂)
	ν ₆	648.2	-7.0	δ(CO ₂)
	ν ₇	490.9	1.6	δ(SO ₂)
SO ₂ OCS	ν ₁	2100.7	0.0	ν(CO)
	ν ₂	1285.9	0.3	ν _a (SO ₂)
	ν ₃	1074.8	2.0	ν _s (SO ₂)
	ν ₄	906.8	5.8	ν(CS)
	ν ₅	497.1	2.2	δ(OCS)
	ν ₆	495.9	1.0	γ(OCS)
	ν ₇	488.8	-0.5	δ(SO ₂)
SO ₂ CS ₂	ν ₁	1636.3	-1.0	ν _a (CS ₂)
	ν ₂	1280.7	-4.9	ν _a (SO ₂)
	ν ₃	1070.8	-2.0	ν _s (SO ₂)
	ν ₄	683.6	-0.5	ν _s (CS ₂)
	ν ₅	489.5	0.2	δ(SO ₂)
	ν ₆	374.5	3.4	δ(CS ₂)
	ν ₇	367.6	-3.5	γ(CS ₂)
SO ₂ N ₂ O	ν ₁	2261.7	18.5	ν(NN)
	ν ₂	1289.7	3.6	ν(NO)
	ν ₃	1283.6	-2.0	ν _a (SO ₂)
	ν ₄	1073.9	1.1	ν _s (SO ₂)
	ν ₅	539.7	1.9	δ(NNO)
	ν ₆	538.9	1.1	γ(NNO)
	ν ₇	490.2	0.9	δ(SO ₂)

hardness²¹ of the electron acceptor, and inversely proportional to the electron affinity,^{20b,22} the mean polarizability^{20c} and the molecular quadrupole moment.^{18,19}

In each case, the datum for the SO₂N₂O adduct is an outlier, which is not consistent with those for the other three complexes. The anomalous position of SO₂N₂O in these relationships has been alluded to above, in connection with the essentially zero charge transfer in this complex, and is related to the differences in the type of intramolecular bonding in N₂O compared with the three carbon-containing molecules, which lead to a different mode of intermolecular binding from that in the remaining three species. The interaction energies of intermolecular complexes may be partitioned into contributions from the attractive charge transfer, electrostatic, polarization and dispersion components and a repulsive Coulomb exchange term, and the net interaction energy can result from a complex interplay among these components. The partitioning into these five terms is clearly quite different for the SO₂N₂O complex from those in the carbon-containing adducts, and an analysis of the decomposition of these interaction energies will be the subject of a future extension of this work.

3.3. Vibrational Spectra

Our observations concerning the computed wavenumber shifts of relatively strongly bound molecular complexes, compared with the wavenumbers of the monomers, have been that these shifts are directly proportional to the interaction energies.²³ In the case of more weakly bound complexes, however, the wavenumber shifts are more erratic, and any dependence on the strength of interaction is determined by the extent to

Table 10 Intermolecular vibrational wavenumbers of the complexes. All modes are of *a* symmetry.

Complex	Mode	Wavenumber/cm ⁻¹	Approximate description ^a
SO ₂ .CO ₂	ν_8	119.7	$\omega(\text{SO}_2)(54) + l(\text{CO}_2)(47)$ (antigeared)
	ν_9	69.5	$\nu(\text{C}\dots\text{O})$
	ν_{10}	51.1	$\omega(\text{SO}_2)(56) + l(\text{CO}_2)(44)$ (geared)
	ν_{11}	29.7	$l(\text{CO}_2)(57)$ (antigeared) + $\text{tw}(\text{SO}_2)(42)$
	ν_{12}	19.5	$\rho(\text{SO}_2)(70) + l(\text{CO}_2)(29)$ (antigeared)
SO ₂ .OCS	ν_8	100.1	$l(\text{OCS})(53)$ (antigeared) + $\omega(\text{SO}_2)(47)$
	ν_9	53.6	$\nu(\text{C}\dots\text{O})$
	ν_{10}	42.5	$\omega(\text{SO}_2)(53) + l(\text{OCS})(46)$ (geared)
	ν_{11}	35.3	$\text{tw}(\text{SO}_2)(61) + l(\text{OCS})(40)$ (antigeared)
	ν_{12}	19.1	$\rho(\text{SO}_2)(77) + l(\text{CO}_2)(23)$ (antigeared)
SO ₂ .CS ₂	ν_8	94.4	$\omega(\text{SO}_2)(95)$
	ν_9	66.6	$\nu(\text{C}\dots\text{O})$
	ν_{10}	53.9	$l(\text{CS}_2)(63)$ (antigeared) + $\omega(\text{SO}_2)(37)$
	ν_{11}	27.1	$\rho(\text{SO}_2)(85)$
	ν_{12}	18.3	$\text{tw}(\text{SO}_2)(88)$
SO ₂ .N ₂ O	ν_8	114.4	$l(\text{NNO})(52)$ (antigeared) + $\omega(\text{SO}_2)(48)$
	ν_9	60.0	$\nu(\text{N}\dots\text{S})$
	ν_{10}	44.6	$l(\text{NNO})(53)$ (geared) + $\omega(\text{SO}_2)(48)$
	ν_{11}	36.4	$l(\text{NNO})(61)$ (antigeared) + $\text{tw}(\text{SO}_2)(39)$
	ν_{12}	15.7	$\rho(\text{SO}_2)(72) + l(\text{NNO})(28)$ (antigeared)

^a Numbers in parentheses are the approximate potential energy distribution contributions of each monomer unit.

which the particular vibrational mode is localized in one monomer or the other.²³ Table 9 shows the intramolecular wavenumbers of the four complexes, with approximate descriptions of the modes, and the shifts relative to the monomers. With the exception of ν_7 of SO₂.OCS, where the SO₂ bending vibration has only an 82 % potential energy distribution, with an 18 % contribution from the OCS bending mode of the OCS sub-molecule, all the modes listed in Table 9 are almost totally localized in one or other of the monomers. The wavenumber shifts of the SO₂ fragments are all less than 5 cm⁻¹, consistent with the weakness of the interactions indicated in Table 8. Since these shifts are all so small, there is no apparent trend in their values, or even their signs, which might be correlated with the interaction energies. The shifts of the modes of the partner molecules are similarly insubstantial, with the exception of the NN stretching mode of N₂O in SO₂.N₂O. Here, too, the signs are irregular, and the spectra of the linear molecules are quite insensitive to the strength of interaction.

The wavenumbers of the intermolecular modes of the complexes are given in Table 10. The intermonomer stretching mode wavenumbers are remarkably constant, at 62 ± 8 cm⁻¹. The hindered rotational vibrations of the SO₂ moiety are coupled to varying degrees with the librational motions of the partner molecules, therefore comparison of the wavenumbers of those modes most closely associated with the SO₂ fragment is not strictly valid. Nevertheless, these mode wavenumbers decrease fairly uniformly in the order wagging > twisting > rocking. Clearly, the perturbations of the spectra of the interacting molecules are minimal, confirming the low interaction energies with which each pair of molecules binds.

4. Conclusions

The molecular complexes formed between sulphur dioxide on one hand, and carbon dioxide, carbonyl sulphide, carbon disulphide and nitrous oxide on the other, each optimize with structures of C₁ symmetry, which are interpreted as being one of a pair of equivalent structures with one SO bond projecting out of the pseudoplane formed by the second SO bond and the

partner molecule. These equivalent structures interconvert through small energy barriers at the C_s positions. The experimentally observed C_{2v} (SO₂.CO₂), C_s (SO₂.OCS⁷ and SO₂.CS₂⁸) and C₁ (SO₂.N₂O⁹) structures may be accommodated by our non-symmetrical models by invoking a tunnelling motion, which is manifested in the experimental rotational spectra of SO₂.N₂O by the splitting of some of the lines. The interaction energies of the complexes involving the carbon-containing molecules are very small, and are proportional to their molecular quadrupole moments, among other properties, while the complex with nitrous oxide exhibits quite anomalous properties. The perturbations of the vibrational spectra of each partner molecule in the complexes are minimal.

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