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AMINO-SUBSTITUTED NAPHTHALENE SULFONIC ACID/GRAPHENE COMPOSITE AS METAL -FREE CATALYSTS FOR OXYGEN REDUCTION REACTIONS

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ABSTRACT. Composites of reduced graphene oxide (rGO) and conducting polymers synthesized from 8-amino-2-naphthalene sulfonic acid (8-ANSA), 5-amino-1-naphthalene sulfonic acid (5-ANSA), 2-amino-1-naphthalene sulfonic acid (2-ANSA), and 4-amino-1-naphthalene sulfonic acid (4-ANSA) were prepared for use as an electrocatalyst in oxygen reduction reaction (ORR). The electrocatalytic activities were examined in oxygen saturated 0.1 M KOH solution using cyclic and linear sweep voltammetry. The best performing polymer composite was found to be GC/poly(8-ANSA)/rGO, with an enhanced electrocatalytic activity over the rGO only and poly(8-ANSA) only films. More than 100 mV positive shift in the onset potential and 1.6 times increase in current density were observed. The catalytic activity of 2-ANSA, 4-ANSA, 5-ANSA, and 8-ANSA was also validated by density functional theory (DFT). Our calculations predict a much better catalytic activity for 8-ANSA than that of 2-ANSA, 4-ANSA, and 5-ANSA.

KEY WORDS: Poly(amino naphthalene sulfonic acid), Reduced graphene oxide, Oxygen reduction reaction, Electrocatalyst

INTRODUCTION

Oxygen reduction reaction (ORR) is one of the most important reactions in a fuel cell, metal-air batteries, chloroalkali electrolyzer, metal corrosion, sensors, industrial electrolytic process and electro-organic reactions [1-3]. However, ORR is kinetically sluggish, involves several steps and many intermediates which makes challenging for researchers in the field. The most efficient catalyst generally considered so far for ORR are noble metals such as platinum (Pt) and its alloy. Nevertheless, these catalysts have a high price, scarce, poor durability, poor tolerance for CO poisoning and methanol crossover. In order to overcome these challenges researchers explore and develop non-precious metals or metal-free catalysts for ORR [1, 4-6].

Graphene is becoming a promising candidate for ORR electrocatalysts due to their excellent conductivity, unique specific surface area, and customized surface chemistry [7, 8]. Graphite oxide has become the most common precursor for the cost-effective and mass production of graphene-based materials. Graphene oxide (GO) needs reduction to restore the disrupted π system/ sp^2 hybridized network using chemical or electrochemical reduction or thermal annealing. The electrochemical reduction method is advantageous over the other in that it is simple, fast, and doesn't require the use of any toxic reducing agents. Reduced graphene oxide (rGO) sheet shows low dispersibility in solvents and tends to restack as a graphite-like structure due to the removal of most of the oxygen functionalities. This aggregation of the rGO sheets reduces the available surface area, limits electron and ion transport which resulted in poor electrochemical performance. One applied strategy to solve this problem is the composite formation of rGO with conducting polymers. The most commonly used conducting polymers are polypyrrole, polyaniline, polythiophene and their derivatives [9-12].

In recent years, graphene composited with poly(3,4-ethylenedioxythiophene): poly(styrenesulfonate) (PEDOT:PSS) [13], polypyrrole [14], polyaniline [15, 16] has been

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reported as an electrocatalyst for ORR. In this work, composite films of amino-substituted naphthalene sulfonic acid based polymers and graphene layers were synthesized for metal-free electrocatalyst for oxygen reduction reaction in alkaline media. Enhanced durability for metal-free over metal catalysts is observed in alkaline media [17].

We previously reported [18] 4-amino-3-hydroxy naphthalene sulfonic acid (AHNSA) for use in catalyzing the oxygen reduction reaction in acidic solution. However, to the best of our knowledge, graphene oxide composited with amino substituted naphthalene sulfonic acids have never been prepared and examined for electrocatalysis in the oxygen reduction reaction.

EXPERIMENTAL

Reagents and apparatus

Sodium nitrate, potassium chloride, nitric acid, 8-amino-2-naphthalene sulfonic acid (8-ANSA), 5-amino-1-naphthalene sulfonic acid (5-ANSA), 2-amino-1-naphthalene sulfonic acid (2-ANSA) and 4-amino-1-naphthalene sulfonic acid (4-ANSA) were purchased from Sigma Aldrich. Sulfuric acid and hydrochloric acid were obtained from Carlo Erba Reagents. Potassium permanganate and potassium hydroxide were obtained from BDH Chemicals Ltd. Hydrogen peroxide, potassium nitrate, potassium hexacyanoferrate(III) and graphite was purchased from Riedel de Haen. Nitrogen and oxygen gases were purchased in cylinders from Chora Gas & Chemical Product Factory.

X-ray diffraction spectra were prepared on MiniFlex X-ray diffractometer using Cu-K α radiation ($\lambda = 0.15406$ nm). PerkinElmer Spectrum 65 FT-IR Spectrometer was used to record the infrared spectra.

A potentiostat/galvanostat (Autolab PGSTAT 128N) connected to a computer with general purpose electrochemical system (GPES v 4.9) and Nova (version 1.9.16) software is used to characterize the electrochemical properties of the electrocatalysts using a rotating disk glassy carbon (RDGC) working electrode (0.196 cm²). Rotating ring-disk electrode (RRDE) measurements were carried out at room temperature on a Pine bipotentiostat (model AFCBP1) connected to a computer with aftermath (version 1.2.5033) software. Glassy carbon disk electrode (0.2472 cm²) surrounded by a Pt ring (0.1859 cm²) were used for RRDE measurements. Silver/silver chloride (Ag/AgCl, 3 M KCl) and a platinum wire served as reference and counter electrodes, respectively. Pine rotation speed controller (AFMSRCE 2957) was used during the rotating disk electrode (RDE) and RRDE measurements.

Synthesis of GO

Graphene oxide was synthesized from graphite powder using modified Hummer's method [19]. Specifically, graphite powder (1 g) and sodium nitrate (0.5 g) were mixed together and added to concentrated sulphuric acid (23 mL) in a beaker under constant stirring at room temperature. The beaker was put in an ice bath for cooling to 0 °C under constant stirring for one hour, and potassium permanganate (3 g) was added gradually to the mixture while keeping the temperature less than 20 °C. The mixture was stirred in a water bath at 35 °C for 6 hours to form a thick paste and the resulting solution was diluted by adding 500 mL of water under vigorous stirring. Subsequently, the suspension was further treated with 30% H₂O₂ solution (5 mL) to terminate the reaction. The resulting mixture was washed with 1:10 HCl aqueous solution followed by filtration. The collected solid material was dried in a vacuum oven to obtain GO.

Electrochemical reduction of GO

Five mg GO was dispersed in 10 mL deionized water and ultrasonicated for 1 h. 20 μ L of the graphene oxide dispersion (0.5 mg mL⁻¹) was drop-casted onto the pre-polished glassy carbon

electrode (GC) or on polymer modified GC (GC/poly(ANSA)) and was left to dry slowly in air. After drying, 3 μL 0.05 wt.% Nafion solution was applied to the surface of the resulting electrode to improve the adhesion of the film to the electrode surface. The electrochemical reduction of GO to rGO was carried out with cyclic voltammetry in the potential range 0 to -1.5 V vs Ag/AgCl at a scan rate of 50 mV s^{-1} in N_2 saturated 0.1 M Na_2SO_4 for 20 cycles [10, 20].

Electropolymerization

Electropolymerization of 2 mM of the monomers (8-ANSA, 2-ANSA, 4-ANSA and 5-ANSA) on bare glassy carbon electrode (GC) or on GC/rGO in 0.1 M HNO_3 was performed using cyclic voltammetry by scanning the potential between -0.8 V and 2.0 V vs Ag/AgCl at a scan rate of 100 mV s^{-1} for 16 cycles. Then, the modified electrode was stabilized for 24 cycles in monomer free 0.5 M H_2SO_4 until a stable cyclic voltammogram was obtained [18].

Electrochemical measurements

The oxygen reduction currents were recorded in oxygen saturated 0.1 M KOH. Cyclic and linear sweep voltammograms were recorded at a scan rate of 20 mV s^{-1} . For the RRDE measurement, the disk and ring currents were recorded as a function of the disk potential scanned between -0.7 V and 0.1 V vs. Ag/AgCl at 10 mV s^{-1} at one rotation speed (1600 rpm). The ring electrode potential was held at +1 V vs. Ag/AgCl [13]. The collection efficiency (N) was determined using 1 M KNO_3 and 10 mM $\text{K}_3\text{Fe}(\text{CN})_6$ deaerated with pure nitrogen. The measured collection efficiency was found to be $38 \pm 1.5\%$.

The oxygen reduction currents were normalized to the geometric area of the electrode and corrected for the background electrode current recorded in pure nitrogen saturated 0.1 M KOH solution [21].

Computational details

All calculations were carried out with the Gaussian 16 program package [22]. The B3LYP 23-25] functional together with the 6-311++G(d,p) basis sets [26] were used. Solvent effects were corrected by using the polarizable continuum model (PCM) in its integral equation formalism [27] together with water as a solvent. Vibrational frequency analyses were performed to verify that the optimized structures are minima without an imaginary frequency and to obtain the free energies of the dimers and their complexes. The natural bond orbital (NBO) analysis [28-29] was used to explore the charges on the atoms before and after the adsorption of the molecular oxygen on the dimers.

RESULTS AND DISCUSSION

The XRD spectra of graphite and GO were recorded in order to verify the structural changes occurring during the conversion of graphite to GO and are shown in Figure 1A. The intense sharp peak ($2\theta = 26.6^\circ$) for graphite powder (Figure 1A(a)) corresponding to the (0 0 2) [30] reflection plane disappeared and shifted to a lower 2θ angle of 11.4° after being oxidized to GO (Figure 1A(b)). The shift is attributed to the intercalation of oxygen-containing groups during oxidation with KMnO_4 [31]. An interlayer distance of ~ 0.78 nm was calculated using Bragg's equation for GO, which is larger than that of graphite powder (~ 0.34 nm). The small peak (Figure 1A(b)) at $2\theta = 42.6^\circ$ is corresponding to the (100) in plane hexagonal atom arrangement [32].

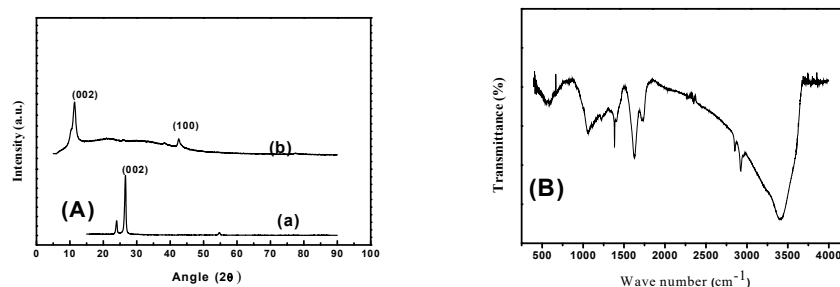


Figure 1. (A) XRD patterns of (a): graphite, (b): GO and (B) FT-IR spectra of GO.

The FTIR spectra of GO (Figure 1B) shows several characteristic peaks. A broad and intense peak at 3414 cm^{-1} attributed to O-H stretching of carboxyl groups, a weak peak at 1735 cm^{-1} assigned to C=O stretching of carbonyl groups, an intense peak at 1632 cm^{-1} corresponds to C=C stretching, a sharp peak at 1399 cm^{-1} and a broad peak at 1065 cm^{-1} are assigned to C-O stretching vibration of carboxyl and alkoxy group, respectively [30, 32, 33].

Figure 2a shows a typical cyclic voltammograms for the electropolymerization of 8-ANSA on a polished glassy carbon rotating disk electrode. During the first cycle, anodic and cathodic peaks at 0.78 V and -0.0811 V , respectively, were observed. In the successive scans, three new anodic peaks appeared at 0.251 V , 0.332 V , and 0.6139 V . Unlike the peak at 0.78 V which is due to the redox activity of the monomer [34], all the peaks increase with increasing number of cycles which shows the formation of polymer film on the surface of the glassy carbon electrode (GC) [18].

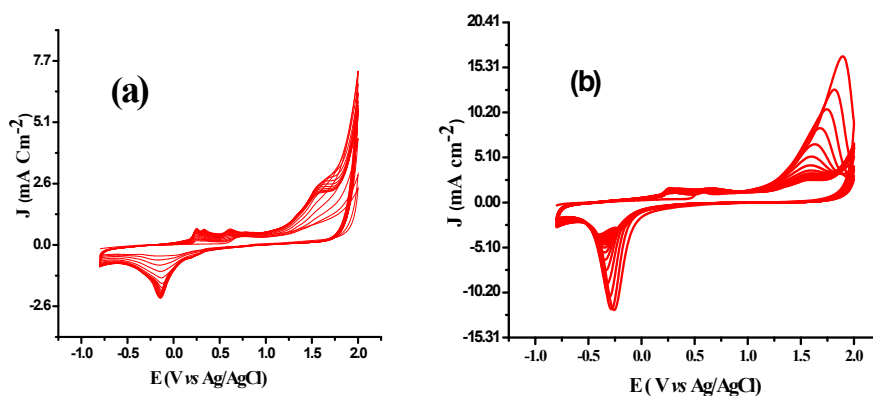


Figure 2. Cyclic voltammograms obtained for the electropolymerization of (a) 8-ANSA on GC and (b) 8-ANSA on GC/rGO from a solution containing 2 mM 8-ANSA monomer in 0.1 M HNO_3 at a scan rate of 100 mV s^{-1} for 16 scans between -0.8 V to 2.0 V vs Ag/AgCl.

Peak currents for the potentiodynamic polymerization of 8-ANSA onto rGO (Figure 2b) coated electrodes are higher than those on the surface of the glassy carbon electrode since modification by rGO leads to a larger electroactive surface area.

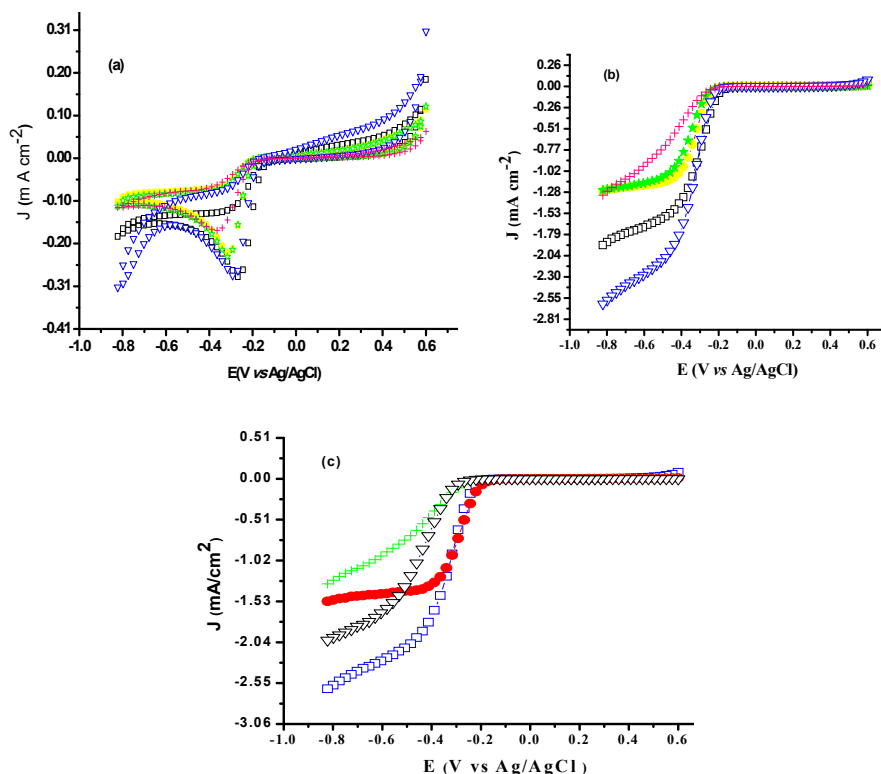


Figure 3. Comparison of the background corrected (a) cyclic voltammograms and (b) linear sweep voltammograms of (+) GC/rGO, (*) GC/poly(5-ANSA)/rGO, (●) GC/poly(4-ANSA)/rGO, (□)GC/poly(2-ANSA)/rGO, and (∇) GC/poly(8-ANSA)/rGO (c) linear sweep voltammograms of (+) GC/rGO, (∇) GC/poly(8-ANSA), (□) GC/poly(8-ANSA)/rGO, and (●) GC/rGO/poly(8-ANSA) electrodes in 0.1 M KOH at a scan rate of 20 mV s^{-1} in 0.1 M KOH. Rotational rate for LSV measurements were 1600 rpm.

After the electrosynthesis of all the monomers (8-ANSA, 2-ANSA, 4-ANSA, and 5-ANSA) on bare RDGC electrode under similar conditions, electrochemical reduction of GO on polymer modified GC electrodes were performed. Cyclic (CV) and linear sweep (LSV) voltammetry measurements were performed in O_2 -saturated 0.1 M KOH to compare the electrocatalytic activity of the four different amino substituted naphthalene sulfonic acid polymers composited with rGO. rGO modified GC electrode was also used in the comparison as a reference. The results are depicted in Figure 3a,b. The ORR activity for all the polymers-rGO modified GC electrode was found to be higher than that of GC/rGO. The onset potentials of GC/poly(2-ANSA)/rGO and GC/poly(8-ANSA)/rGO were more positive compared to that of GC/poly(4-ANSA)/rGO and GC/poly(5-ANSA)/rGO composite films. In terms of the current density, GC/poly(8-ANSA)/rGO has the highest value among all the polymers. Thus, GC/poly(8-ANSA)/rGO composite was chosen for further analysis.

The reaction mechanism of oxygen reduction involves diffusion in the bulk of the electrolyte and the conducting films, adsorption process, and electron transfer [35]. Any of these processes might be rate controlling and can be one of the criteria for the catalytic activity of materials. To

explain the catalytic activity of poly(5-ANSA), poly(4-ANSA), poly(2-ANSA), and poly(8-ANSA), quantum-chemical calculations of the electronic structure for all the amino substituted naphthalene sulfonic acid and their adsorption complexes with molecular oxygen were performed using density functional theory (DFT). The dimers (Figure 4) and their adsorption complexes with oxygen were used for all the calculations. To identify the adsorption sites, we used the natural charges from the natural bond orbital (NBO) analysis. From the results in Table 1, the maximum negative charge are concentrated on carbon atoms in positions C₃, C₄, C₈, C₉, C₁₄, C₁₅, C₁₇, and C₁₈. Each carbon atoms can in principle be an adsorption site for oxygen binding, however, the carbon atoms with larger negative charges are more preferable for the oxygen adsorption. The results listed in Table 1 show a clear increase in the negative charge on O₂ (from 0.0 to -0.367; see the last two rows of Table 1), and a positive charge on the adsorption sites (C₈-C₉) in the complexes (2-ANSA-O₂, 4-ANSA-O₂, 5-ANSA-O₂, and 8-ANSA-O₂), indicating the adsorption of O₂ on the dimers.

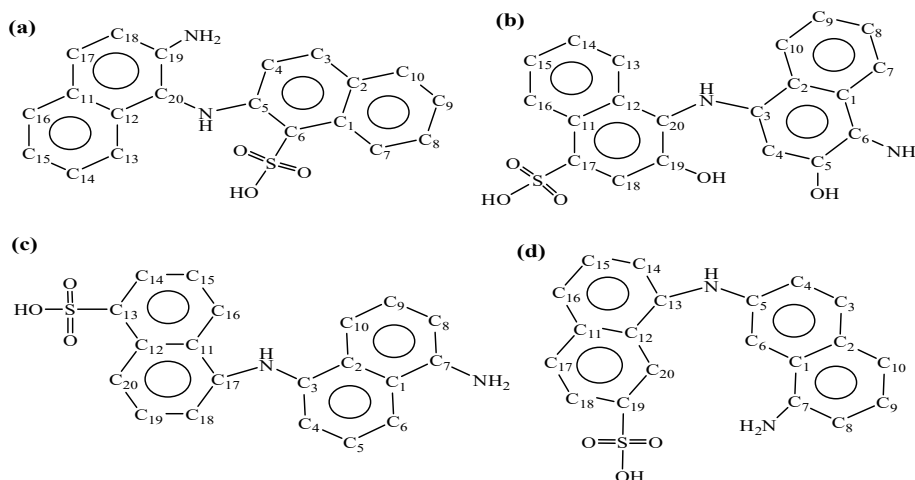


Figure 4. Dimers of (a) 2-ANSA, (b) 4-ANSA, (c) 5-ANSA, and (d) 8-ANSA molecular cluster. Hydrogen atoms were omitted for clarity.

The Gibbs free energy of formation (ΔG_f) (Table 2) between the oxygen atoms and the carbon atoms in the positions 8 and 9 are lower for all ANSA-O₂ complexes than for other carbon atoms (e.g. C₁₄ and C₁₅). This clearly shows that positions 8 and 9 are the most preferable O₂ adsorption sites.

The electron density transferred to O₂ and occupies the antibonding MO's of O₂ (adsorbed) which is responsible for the destabilization of the O-O bond in O₂. This is evident from the O-O bond length which is elongated to 1.498, 1.5, 1.497 and 1.492 Å for 2-ANSA-O₂, 4-ANSA-O₂, 5-ANSA-O₂, and 8-ANSA-O₂, respectively, from that of 1.21 Å for the bare O₂ molecule. The O-C_x and O-C_y bond lengths also show the adsorption of the O₂ molecule. Thus, chemisorbed O₂ molecules have a fairly high degree of activation and can be easily reduced, which accounts for the catalytic activity of the ANSA polymers. Furthermore, 8-ANSA followed by 2-ANSA, with lowest ΔG_f and more possible adsorption sites, are the best electrocatalytic materials for oxygen reduction reaction. This is in agreement with the experimental observation (Figure 3a,b).

Table 1. Natural charges obtained from the natural bond orbital (NBO) analysis in the dimers and dimer-O₂ complexes (where the binding site in all dimers is C₈-C₉), calculated using B3LYP/6-311++G(d,p)/PCM/water calculations. Other possible binding sites are also shown in **bold-italic** fonts.

Atom	2-ANSA	2-ANSA-O ₂	4-ANSA	4-ANSA-O ₂	5-ANSA	5-ANSA-O ₂	8-ANSA	8ANSA-O ₂
C ₁	-0.047	-0.060	-0.065	-0.030	-0.081	-0.049	-0.084	-0.089
C ₂	-0.097	-0.073	-0.085	-0.051	-0.060	-0.029	-0.066	-0.102
C ₃	-0.123	-0.123	0.175	0.186	0.182	0.190	-0.165	-0.211
C ₄	-0.232	-0.219	-0.301	-0.294	-0.240	-0.236	-0.227	-0.198
C ₅	0.239	0.228	0.280	0.284	-0.202	-0.201	0.155	0.106
C ₆	-0.370	-0.333	0.117	0.088	-0.202	-0.234	-0.217	-0.202
C ₇	-0.152	-0.235	-0.187	-0.234	0.189	0.139	0.177	0.244
C ₈	-0.207	0.077	-0.219	0.071	-0.270	0.077	-0.252	0.098
C ₉	-0.222	0.072	-0.201	0.096	-0.186	0.078	-0.207	0.081
C ₁₀	-0.168	-0.122	-0.202	-0.164	-0.248	-0.213	-0.226	-0.205
C ₁₁	-0.079	-0.082	-0.072	-0.075	-0.071	-0.077	-0.070	-0.068
C ₁₂	-0.050	-0.051	-0.051	-0.051	-0.043	-0.046	0.013	0.009
C ₁₃	-0.208	-0.208	-0.201	-0.193	-0.265	-0.260	0.185	0.192
C ₁₄	-0.195	-0.193	-0.170	-0.183	-0.174	-0.190	-0.243	-0.254
C ₁₅	-0.236	-0.234	-0.204	-0.202	-0.206	-0.207	-0.173	-0.170
C ₁₆	-0.162	-0.161	-0.191	-0.176	-0.149	-0.137	-0.217	-0.225
C ₁₇	-0.170	-0.140	-0.231	-0.231	0.205	0.202	-0.147	-0.148
C ₁₈	-0.217	-0.237	-0.205	-0.207	-0.250	-0.255	-0.205	-0.206
C ₁₉	0.172	0.180	0.318	0.322	-0.169	-0.168	-0.302	-0.304
C ₂₀	0.077	0.068	0.120	0.116	-0.224	-0.231	-0.175	-0.175
O ₄₃	-	-0.327	-	-0.336	-	-0.329	-	-0.345
O ₄₄	-	-0.335	-	-0.331	-	-0.322	-	-0.367

Table 2. Change in Gibbs free energy of formation, ΔG_f (kcal/mol), equilibrium interatomic distances (Å) between the oxygen atoms, $r(\text{O}=\text{O})$, between the oxygen atoms and the carbon atoms, $r(\text{O}-\text{C})$, in the dimers and dimer-O₂ complexes.

	Binding site (C _x -C _y)	ΔG_f	$r(\text{O}=\text{O})$	$r(\text{O}-\text{C}_x)$	$r(\text{O}-\text{C}_y)$
O ₂	-	-	1.205	-	-
2ANSA-O ₂	C ₈ -C ₉	10.73	1.498	1.490	1.467
	C ₁₄ -C ₁₅	10.89	1.499	1.495	1.467
4ANSA-O ₂	C ₈ -C ₉	12.99	1.500	1.489	1.467
	C ₁₄ -C ₁₅	14.94	1.497	1.484	1.466
5ANSA-O ₂	C ₈ -C ₉	14.90	1.497	1.484	1.466
	C ₁₄ -C ₁₅	12.68	1.497	1.484	1.466
8ANSA-O ₂	C ₈ -C ₉	10.71	1.492	1.545	1.455
	C ₁₄ -C ₁₅	11.14	1.497	1.535	1.455

The electrocatalytic activities for ORR of GC/rGO, GC/poly(8-ANSA), GC/poly(8-ANSA)/rGO and GC/rGO/poly(8-ANSA) were studied using linear sweep voltammetry (LSV) in N₂ and O₂ saturated 0.1 M KOH solution and compared with each other in Figure 3c. Background corrected polarization curves (Figure 3c) at 1600 rpm shows GC/poly(8-ANSA)/rGO has the highest current density and more positive onset potential than GC/poly(8-ANSA) as well as GC/rGO, which demonstrate the good electrocatalytic activity of GC/poly(8-ANSA)/rGO. GC/rGO/poly(8-ANSA) composite exhibits lower electrocatalytic activity towards ORR. At -0.6 V for instance, the current density recorded for ORR on the GC/poly(8-ANSA)/rGO is 2.27 mA cm⁻² which is much higher than GC/rGO/poly(8-ANSA) (1.44 mA cm⁻²). This is attributed to oxidation potential applied to electropolymerized 8-ANSA on

GC/rGO to produce GC/rGO/ poly(8-ANSA) which could also possibly oxidize the previously synthesized rGO forming less conductive graphene oxide (GO).

The kinetics of the ORR activity of GC/poly(8-ANSA)/rGO were further investigated at different electrode rotational rates from 400 to 2500 rpm in 0.1 M KOH as shown in Figure 5a. The onset of the O₂ reduction wave on GC/poly(8-ANSA)/rGO begins at a potential close to -0.21 V (*vs* Ag/AgCl). The current density increased when the rotational rates were increased from 400 rpm to 2500 rpm, indicating that the ORR is controlled by mass diffusion [13].

Koutecky-Levich (K-L) equation was applied in mixed kinetic -diffusion controlled region in order to estimate the transferred electron number per oxygen molecule involved in the ORR. The K-L equation can be represented as follows:

$$\frac{1}{J} = \frac{1}{J_L} + \frac{1}{J_k} = \frac{1}{0.2nF C_0 D_0^{2/3} \nu^{-1/6} \omega^{1/2}} + \frac{1}{nF k_{O_2} C_{O_2} \Gamma} = \frac{1}{J_L} + \frac{1}{nF k_f C_{O_2}} \quad (1)$$

where J is the measured current density, J_k and J_L are the kinetic and diffusion-limiting current densities, ω is the angular velocity of the disk (in rpm), n is the overall number of electrons transferred in oxygen reduction, F is the Faraday constant ($F = 96485 \text{ C mol}^{-1}$), C_0 is the bulk concentration of O₂ ($1.2 \times 10^{-3} \text{ M}$), D_0 is the diffusion coefficient of oxygen ($1.9 \times 10^{-5} \text{ cm}^2 \text{ s}^{-1}$), ν is the kinematic viscosity of the electrolyte ($\nu = 0.01 \text{ cm}^2 \text{ s}^{-1}$), k_{O_2} is the electron transfer rate constant, k_f is the heterogeneous rate constant, and Γ is the surface coverage of the catalyst on GC electrode as given in Table 3. The surface coverage Γ can be evaluated from the equation $\Gamma = Q/nFA$, where Q is the charge obtained by integrating the cathodic peak under the background correction at low scan rate from Figure 3a [2, 36].

K-L plots (J^{-1} *vs* $\omega^{-1/2}$) for GC/poly(8-ANSA)/rGO (Figure 5b) are parallel and linear, confirming that ORR follows first-order kinetics with respect to O₂ molecule. Furthermore, the plots do not pass through the origin indicating a mixed kinetic -diffusion controlled mechanism [15, 29]. The number of electrons transferred (n) and J_k can be obtained from the slope and intercept of the K-L plots [30]. The heterogeneous rate constants (k_f) for the reduction of oxygen were calculated from J_k and reported in Table 3. The ORR occurs either via direct four electrons pathway where O₂ is reduced to OH^- or 2-electron reduction pathway, where O₂ is reduced to HO_2^- [4]. The electron transfer number were calculated using K-L plots over the potential region of -0.36 to -0.43 V to be 2.8 - 2.5 for GC/poly(8-ANSA)/rGO, 2-2.13 for GC/rGO, and 1.8-2.5 for GC/poly(8-ANSA). These result demonstrated that the ORR process at the GC/poly(8-ANSA) and GC/rGO composite proceed with two electrons ($2e^-$) pathway with the formation of HO_2^- but a mixture of HO_2^- and OH^- for GC/poly(8-ANSA)/rGO [4]. The direct, green and in situ electroreduction of O₂ dissolved in alkaline media to HO_2^- on GC/rGO and GC/poly(8-ANSA) electrodes is an attractive alternative to the current anthraquinone-based industrial production of H₂O₂, which produces a large number of chemical pollutants in a multistep energy intensive processes [39].

The RRDE was used further to evaluate the ORR pathway by monitoring the intermediate peroxide species of HO_2^- generated at the disk electrode in the alkaline medium during the ORR process at the GC/poly(8-ANSA)/rGO, GC/rGO, and GC/poly(8-ANSA) electrodes. As shown in Figure 5d, all three catalysts generate ring currents over the potential range from -0.2 to -0.7 V for ORR. The electron transfer numbers (n) derived from RRDE result using equation 2 [5] over the potential range of -0.36 to -0.43 V to be 2.1 - 2.18 for GC/poly(8-ANSA), 2.14 - 2.15 for GC/rGO, and 2.76 - 2.82 for GC/poly(8-ANSA)/rGO, being consistent with the RDE results. Hence, the HO_2^- yields determined using equation 3 [5] for GC/poly(8-ANSA)/rGO (57%) significantly lower than GC/rGO (93%), and GC/poly(8-ANSA) (92%). All these values demonstrate the synergistically enhanced electrocatalytic efficiency of the composite (GC/poly(8-ANSA)/rGO).

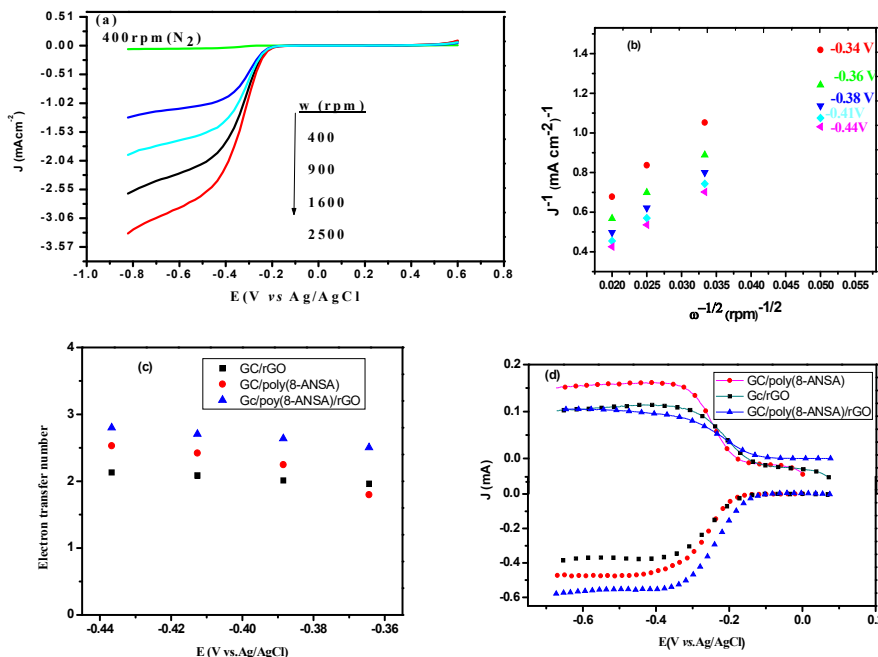


Figure 5. (a): Linear sweep voltammograms of GC/poly(8-ANSA)/rGO at different rotation speeds at a scan rate of 20 mV s⁻¹, (b): K-L plots (J^{-1} vs $\omega^{-1/2}$) at different electrode potentials in 0.1 M KOH, (c): Electron transfer number as a function of potential of GC/rGO, GC/poly(8-ANSA), and GC/poly(8-ANSA)/rGO, and (d) RRDE voltammograms of GC/rGO, GC/poly(8-ANSA), and GC/poly(8-ANSA)/rGO at 1600 rpm. The ring current was polarized at +1.0 V.

$$n = \frac{4i_d}{i_d + \frac{i_r}{N}} \quad (2)$$

$$\% \text{HO}_2^- = \frac{200i_r}{i_d + \frac{i_r}{N}} \quad (3)$$

$$N = -\frac{i_r}{i_d} \quad (4)$$

where N is the collection efficiency of the rotation ring disk electrode (RRDE), and i_d and i_r are the disk and ring electrode currents, respectively.

The logarithm of J_k was plotted against the potential to analyze the Tafel behavior for GC/poly(8-ANSA)/rGO, GC/poly(8-ANSA), and GC/rGO catalysts. The reaction kinetics can be evaluated using Tafel equation (5) [40]:

$$\log J_k = -\frac{1}{b}(E - E^0) + \log J_0 \quad (5)$$

$$b = \frac{2.303RT}{\alpha nF} \dots \quad (6)$$

where E is the measured potential of the working electrode, E^0 is the equilibrium open-circuit potential, b is the Tafel slope, R is gas constant (8.314 J mol⁻¹ K⁻¹); T is the Kelvin temperature (298.15 K), J_k is the kinetics current density, and J_0 is the exchange current density.

Using equation 5, the high value of J_0 and low value of b will give high J_k value under same overpotential ($E-E_0$). J_0 is a key parameter which reflects intrinsic rates of electron transfer between the solution and the electrode. It can be used to judge the catalytic efficiency of materials [40, 41].

The Tafel plots clearly show the electrocatalytic activity differences among the electrode materials. At -0.36 V, for instance, the kinetic current density of ORR on the GC/poly(8-ANSA)/rGO is 7.53 mA cm^{-2} , which is significantly higher than GC/rGO (0.58 mA cm^{-2}) and GC/poly(8-ANSA) (0.82 mA cm^{-2}). Tafel slopes and exchange current density were estimated from the linear portion of the Tafel plot at low current density region (Figure 6a) and presented in Table 3. Therefore, incorporation of 8-ANSA between GC and rGO reduces the Tafel slope from 74 mV dec^{-1} to 68 mV dec^{-1} indicating more favorable kinetics for ORR [42]. The exchange current densities (J_0) were estimated from the intercept of a linear portion of the Tafel plots [41]. The J_0 values for GC/8-ANSA/rGO much higher than that of GC/rGO and GC/8-ANSA showing an improved catalytic efficiency of the composite for ORR.

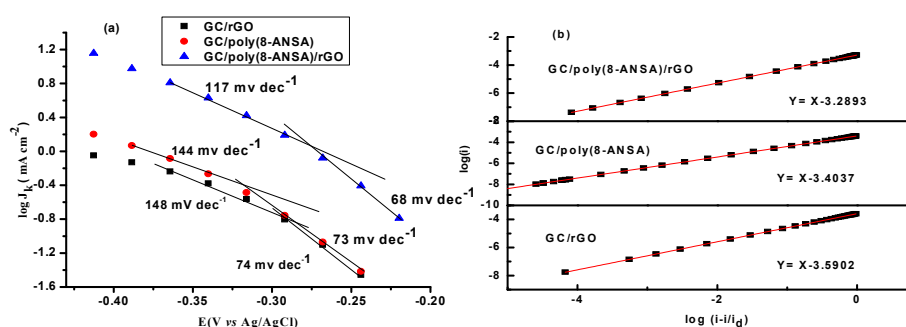


Figure 6. (a) Tafel plots for the ORR at GC/poly(8 ANSA)/rGO (▲), GC/poly(8-ANSA) (◆), and GC/rGO (■) at an electrode rotation rate of 1600 rpm. (b) reaction order plot for oxygen reduction reaction on GC/rGO, GC/poly(8-ANSA), and GC/poly(8-ANSA)/rGO (the slope stands for the reaction order number).

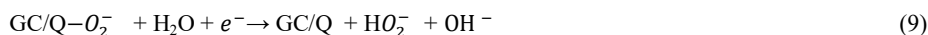
The reaction order (m) of ORR on GC/rGO, GC/poly(8-ANSA), and GC/poly(8-ANSA)/rGO catalysts was checked further by plotting $\log i$ versus $\log[(i-i_d)]$ [43]. From Figure 6b and Table 3, the first order dependence of the kinetics of the ORR was performed on GC/rGO, GC/poly(8-ANSA), and GC/poly(8-ANSA)/rGO.

Table 3. Tafel slopes and exchange current densities for the ORR at GC/rGO, GC/8-ANSA, and GC/8-ANSA/rGO.

Catalyst	m	n		α	b (mV dec ⁻¹)	J_0 (mA cm ⁻²)	Γ (mol cm ⁻²)	k_f (cm s ⁻¹)
		K-L	RRDE					
GC/rGO	1	2.04	2.01	0.8	74	1.76×10^{-5}	1.87×10^{-8}	2.7×10^{-3}
GC/8-ANSA	1	2.3	2.12	0.81	73	1.78×10^{-5}	2.01×10^{-8}	3.8×10^{-3}
GC/8-NSA/rGO	1	2.7	2.8	0.87	68	9.58×10^{-5}	2.36×10^{-8}	3.5×10^{-2}

The value of k_f depends on the intrinsic activity of the catalyst sites and the surface concentration of catalyst sites [37]. In the present work, the surface coverage of all the catalyst films have the same order of magnitude ($10^{-8} \text{ mol cm}^{-2}$) therefore the surface concentration of the catalysts assumed to be constant. GC/rGO, GC/poly(8-ANSA), and GC/poly(8-ANSA)/rGO yielded a slope of 74, 73, and 68 mV dec^{-1} , respectively, at low current density and the Tafel

slopes increase at a higher current density. A Tafel slope of 60 mV dec^{-1} at 25°C would suggest a fast electron transfer reaction followed by a rate-determining chemical step. Furthermore, an increase in Tafel slope at more negative potential in alkaline solution would also attribute to the chemical rate-determining step [37, 44]. A mechanistic scheme that is consistent with these results is as follows:



where, GC/Q is glassy carbon modified with rGO, poly(8-ANSA) or poly(8-ANSA)/rGO. Reaction 8 is the rate determining slow chemical step occurred after a fast electron transfer step (reaction 7).

The specific performance of GC/poly(8-ANSA)/rGO is compared with similar other values reported in the literature in Table 4.

Table 4. Comparison of materials and specific performance of the catalyst in 0.1 M KOH.

Catalyst	Peak potential (vs. Ag/AgCl) (V)	b (mv dec^{-1})	α	J^0 (mA cm^{-2})	Reference
rGO-DAB	-0.28	62	-	-	5
PEDOT:PSS/rGO	-0.27	-	-	-	13
Py-EGO	-0.28	-	-	-	7
Fe-N/ rGOHN500	-	88	0.67	2×10^{-5}	31
Fe-N/rGOsonic	-	86	0.69	2×10^{-5}	31
(G-dye-FeP)n MOF	-0.23	-	-	-	36
CNx/graphene	-0.19	-	-	-	45
PPy/rGO	-0.30	-	-	-	45
N -doped graphene	-0.32	-	-	-	46
GC/poly(8ANSA)/rGO	-0.29	68	0.87	9.58×10^{-5}	This work

-DAB-1,4-diaminobutane, -Py-EGO-pyridine functionalized graphene nanosheets, -HN500- thermal annealed under $\text{H}_2\text{-N}_2$ atmosphere at 500°C , -Sonic-ultrasonication, -(G-dye-FeP)n MOF-dye functionalized rGO-iron-porphyrin metal-organic framework, -Carbon-nitrogen ($x = 0.15$), -PPy -polypyrrole.

CONCLUSION

We applied quantum mechanical (QM) calculations to evaluate the catalytic activity of poly(8-amino-2-naphthalene sulfonic acid), poly(2-amino-1-naphthalene sulfonic acid), poly(5-amino-1-naphthalene sulfonic acid), and poly(4-amino-1-naphthalene sulfonic acid). Our investigations have shown that the existence of electrocatalytic activity for oxygen reduction reaction on different amino substituted naphthalene sulfonic acid. QM calculations predict a much better catalytic activity for poly(8-amino-2-naphthalene sulfonic acid). The preparation and characterization of 8-ANSA, 2-ANSA, 5-ANSA, and 4-ANSA with rGO as composite electrocatalytic material for oxygen reduction reaction were investigated electrochemically. The polymerization of all monomer and conversion of GO to rGO was done using cyclic voltammetry. The composites of poly(8-ANSA) and rGO was found as a promising material for oxygen reduction reaction due to its better electrocatalytic performance. Analysis of the electrochemical reduction of oxygen at GC/poly(8-ANSA)/rGO exhibits enhanced catalytic activity with high exchange current density and high kinetic current towards oxygen reduction reaction in alkaline solution compared to GC/rGO and GC/poly(8-ANSA). GC/rGO and

GC/poly(8-ANSA) have good applicability in catalyzing the reduction of oxygen for in situ production of H₂O₂.

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