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## Donor-substituted phosphanes – surprisingly weak Lewis donors for phosphonium cation stabilisation

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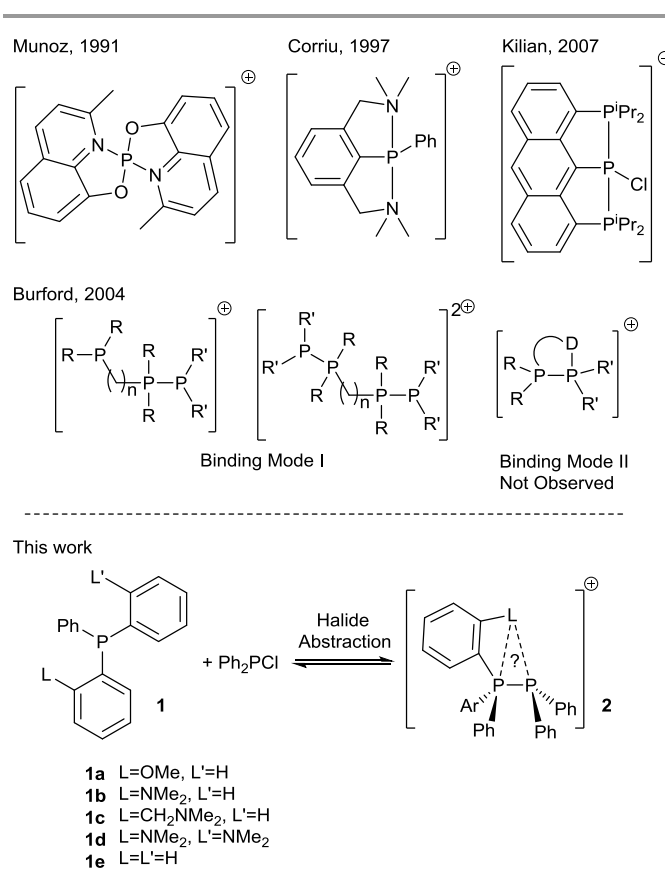
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**Paradoxically, N- and O- donor substituted tri-arylphosphanes are shown to be weaker donors than PPh<sub>3</sub> when binding the soft Lewis acid moiety [PPH<sub>2</sub>]<sup>+</sup>. This arises from internal solvation and rehybridisation at phosphorus, precluding chelation and increasing steric demand, in direct contrast to coordination modes observed for metal complexes.**

Phosphacations in the form of highly electron poor, four coordinate P(V) cations have been shown to exhibit high Lewis acidity enabling C-F bond activation<sup>1</sup> and their use as catalysts for hydrogenation *via* frustrated Lewis pair (FLP) chemistry has been demonstrated.<sup>2,3</sup> In contrast, two coordinate, cationic P(III) (phosphenium) species are less well studied.<sup>4</sup> These highly reactive species rapidly insert into C-X bonds,<sup>5,6</sup> and in general require significant steric bulk and strongly  $\pi$ -donating substituents to render them sufficiently stable for isolation.<sup>4</sup> Stability may also be imparted by quenching the Lewis acidity with a  $\sigma$  donor (Lewis basic) species,<sup>7</sup> regenerating the three coordinate phosphorus centre, which for chelating ligands may permit higher coordination modes (Figure 1).<sup>8-10</sup> Phosphanes themselves are good donors for this and the stability of phosphane-phosphenium complexes is attributed to the thermodynamic favourability of the P-P bond.<sup>11</sup> Interestingly, whilst tetracoordinate P(III) monocations are known with internal chelating ligands,<sup>12,13</sup> there are very few examples for phosphorus donors or intramolecular chelates (Figure 1). Burford's attempted syntheses using symmetrical diphosphane ligands instead gave either phosphane-phosphenium complexes with three coordinate phosphacation centres and a free, unbound donor centre or symmetrical bis-phosphenium species for alkyl bridged ligands,<sup>14</sup> or rearrangement in the case of aryl bridged species.<sup>15</sup>



**Figure 1:** Previously reported and proposed phosphane-phosphenium coordination modes with chelating and multidentate ligands

We hypothesised that combination of a phosphane donor moiety with rigidly linked first row main-group donors would favour the formation of the elusive binding mode II by preorganisation towards binding and the increased stability of hypervalent bonding involving more electronegative elements. In this report, we describe the synthesis of a family of simple donor functionalised phosphane-derived phosphane-phosphenium salts; the effect of donor substitution on the

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Electronic Supplementary Information (ESI) available: [details of any supplementary information available should be included here]. See DOI: 10.1039/x0xx00000x

overall donor strength and resultant cation stability is discussed.

## Results and discussion

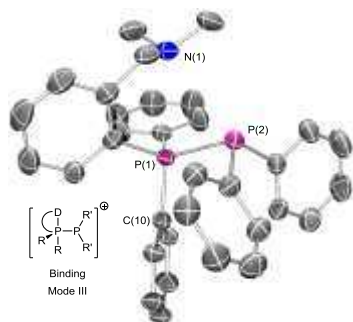
### Initial Synthesis and Structure

No reaction is observed on combination of phosphanes **1a** to **1e** with Ph<sub>2</sub>PCl in DCM; addition of one equivalent of NaBAR<sup>F</sup> lead in all cases to complete loss of <sup>31</sup>P NMR signals for both starting materials and signals indicative of phosphane-phosphenium formation – see Table 1. In contrast to the broad, unresolved signals often seen for **2e** salts,<sup>15</sup> <sup>1</sup>J<sub>P,P</sub> coupling could be resolved in solution for **2b**[BAR<sup>F</sup>], **2c**[BAR<sup>F</sup>] and **2d**[BAR<sup>F</sup>] (Table 1).

	BAR <sup>F</sup>	OTf	GaCl <sub>4</sub>
<b>[2a]</b>	16.4, -6.4 -	17.1 -6.4 (355)	18.9, -5.8 (307)
<b>[2b]</b>	13.7, -8.4 (345)	13.8, -8.9 (344)	13.8, -8.8 (344)
<b>[2c]</b>	17.0, -6.1 (335)	-, - <sup>a</sup> -	17.0, -6.5 <sup>a</sup> (332)
<b>[2d]</b>	12.8, -6.3 (335)	12.8, -6.8 (331)	12.8, -6.6 (335)
<b>[2e]</b>	13.7, -10.1 -	15.3, -10.6 (~340)	13, -13 <sup>15</sup> (340)

**Table 1:** <sup>31</sup>P NMR Parameters for phosphane-phosphenium salts in DCM, <sup>1</sup>J<sub>P,P</sub> / Hz in parenthesis.<sup>a</sup> Not cleanly formed.

The reaction with Me<sub>3</sub>SiOTf or GaCl<sub>3</sub> likewise resulted in clean chloride abstraction and the quantitative formation of the desired triflate or [GaCl<sub>4</sub>]<sup>-</sup> salts except for **1c** which gave complex mixtures for both reagents. Significant variation is seen for the <sup>1</sup>J<sub>P,P</sub> coupling constants, especially for **[2a]** salts indicating varying degree of anion association. Attempted synthesis of the cheaper [AlCl<sub>4</sub>]<sup>-</sup> salts by halide abstraction with AlCl<sub>3</sub> lead to complex behaviour with multiple species present in solution by <sup>31</sup>P NMR, likely due to the more coordinating nature of the anion coupled with competition from the harder nitrogen donor centres. Both the <sup>31</sup>P chemical shifts and coupling constants are comparable to those seen for **2e**, implying that binding mode II is not adopted. This was confirmed upon successful isolation and characterisation of single crystals of **[2c]BAR<sup>F</sup>**, revealing that it exhibits the unexpected mode III with short N(1)-P(1) contacts. (Figure 2)

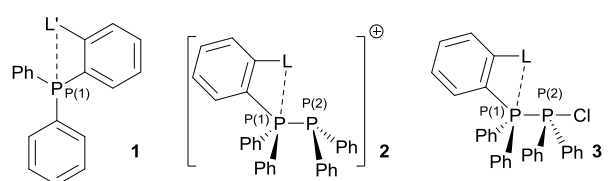


**Figure 2:** Crystal structure of **[2c]BAR<sup>F</sup>**. Drawn with ellipsoids at 50% probability; hydrogens and disordered anion omitted for clarity.

Compound **[2c]BAR<sup>F</sup>** crystallises with a single ion pair in the asymmetric unit, with disordered CF<sub>3</sub> units in the counteranion (Figure 2). The P(1)-P(2) bond length is unexceptional but slightly long at 2.2477(9) Å (c.f. 2.2302(13) Å for **[2e]OTf**),<sup>15</sup> and there is a close contact between the donor group and the adjacent phosphorus centre (d(N(1)-P(1)) = 3.014(3) Å, less than the sum of the van der Waals radii (3.35 Å)<sup>16</sup>) with the nitrogen lone-pair clearly oriented towards the phosphorus centre. The donor phosphorus may therefore be described as either a monocapped tetrahedron or a highly distorted trigonal bipyramid – the sum of equatorial angles = 334.37 ° (c.f. 328.4 ° for a tetrahedron) but the axial N(1)-P(1)-C(10) angle is 175.4(1) ° suggestive of trigonal bipyramid geometry. This is therefore the rather unusual situation in which a P(III) centre is acting as a Lewis acid and a Lewis base simultaneously. Similar internal coordination has been seen for internally solvated phosphanes,<sup>17-20</sup> and for P-alkylated phosphonium salts bearing rigid bearing 2-donor-1-naphthyl fragments<sup>17,18</sup> (the more flexible N,N-dimethylbenzylamino- substituted variants instead react with alkylating agents at the harder nitrogen centre.) Despite the inequivalence of the P(1)-bound phenyl and N-methyl groups in the crystal structure, these give single resonances in the <sup>1</sup>H NMR down to -50 °C, indicating a dynamic structure in solution. Notably, this differs significantly from the chelating P,N coordination mode observed for **1c** with transition metal Lewis acids.<sup>21</sup> The remaining donor-functionalised salts were found to decompose on exposure to hexane or Et<sub>2</sub>O (See ESI), implying that the additional donor moieties destabilise the cation with respect to the easily isolated **[2e]** and, with the coupling constant variation, that there is significant anion dependence on their stability.

### Computational Investigation and VT NMR Studies

Computational methods have been used to explore and support the crystallographic observations of internal solvation. (See ESI for further details)



	Bond Lengths / Å			Fuzzy Bond Indices		
	N(1)-P(1)	P(1)-P(2)	P(2)-Cl(1)	N(1)-P(1)	P(1)-P(2)	P(2)-Cl(1)
<b>1b</b>	2.92104			0.1445		
<b>1c</b>	2.89306			0.1912		
<b>1e</b>						
Ph <sub>2</sub> PCl			2.12225			1.3485
<b>2b</b>	2.92143	2.26457		0.1311	1.0217	
<b>2c</b>	2.99666	2.27408		0.1440	1.0167	
<b>2e</b>		2.25743			1.0459	
<b>3b</b>	3.07198	3.21079	2.16476	0.0907	0.2718	1.2758
<b>3c</b>	3.01817	2.31104	3.08682	0.1362	0.9761	0.3831
<b>3e</b>		3.47776	2.15110		0.1847	1.3031

**Table 2:** Bond Parameters (calculated at the M06-2X/6-311g+(d,p) level).

	NBO Charges						
	1b	1c	1e	Ph <sub>2</sub> PCl	2b	2c	2e
P(1)	0.858	0.871	0.840		1.313	1.325	1.284
P(2)				0.891	0.725	0.719	0.710
N(1)	-0.587	-0.586			-0.606	-0.596	
C(1)	-0.348	-0.348	-0.324		-0.402	-0.429	-0.410 <sup>a</sup>
Cl(1)				-0.344			

**Table 3:** Computational NBO Atomic Charge Distribution (calculated at the M06-2X/6-311g+(d,p) level), C(1) is the apical carbon for the hypervalent bond<sup>a</sup> mean value, see ESI..

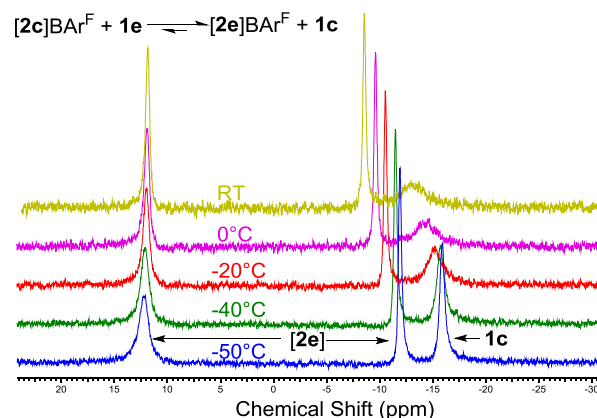
Calculations were performed on full, rather than model structures, given the likely significant steric and electronic influence of the adjacent  $\pi$  systems. In light of this computational demand, we selected three cations – **2b,c,e** – for modelling and also the hypothetical adducts **3b,c,e** for comparison of donor strength in a neutral complex (Table 2). The optimised structures of **1b**, and **2c** were in agreement with experimental data,<sup>22</sup> giving confidence in the model. The computed structures for **2b** also exhibited a short N(1)-P(1) contact, again in contrast to the behaviour of the ligand **1b** with transition metal Lewis acids, but similar to the more substituted analogues.<sup>18,19</sup> Examination of the computed molecular orbitals show that the HOMOs of **1b**, **1c** and **1e** all correspond to phosphorus centred lone pairs, with energies of -7.18 eV, -7.16 eV and -7.46 eV respectively; internal coordination therefore significantly raises the energies of the lone pairs (expected to increase donor strength) whilst also increasing positive charge at phosphorus. Furthermore, the HOMO-1 for **1b** and **1c** are in both cases P-N  $\sigma$  bonding interactions. For the cations, the donor-functionalised species are enthalpically stabilised relative to **2e**, (Table 4) though only to a small degree. In terms of free energy, **2c** is in fact slightly destabilised with respect to **2e**, presumably reflecting the entropic cost of binding the otherwise freely rotating benzyl moiety.

	$1 + 2e \longrightarrow 2 + 1e$		$1 + Ph_2PCl \rightleftharpoons 3$	
	$\Delta E^b /$ kcalmol <sup>-1</sup>	$\Delta G^{a,b} /$ kcalmol <sup>-1</sup>	$\Delta E /$ kcalmol <sup>-1</sup>	$\Delta G^a /$ kcalmol <sup>-1</sup>
<b>1b</b>	-1.5	-0.2	-11.3	4.9
<b>1c</b>	-0.9	1.3	-9.4	8.6
<b>1e</b>	0	0	-10.6	4.1

**Table 4:** Computational Energetics (calculated at the M06-2X/6-311g+(d,p) level). a. Zero point corrected at 298.15K b. stabilisation relative to **2e**.

The calculated Mayer's Fuzzy Bond Indices show a significant degree of covalent bonding between the nitrogen and phosphorus centres for both **1b** and **1c**, which decreases in **2b** and **2c** compared to the free phosphanes. A slight increase in negative charge at the donor nitrogen and apical carbon centre is seen in the cationic complexes, coupled with an increased positive charge at P(1) (Table 3) relative to **2e**. This suggests a decrease in P-N bonding, supported by P-N bond elongation on complex formation and attributed to steric repulsion. With this in mind, the increased exothermicity of P-

P bonding with respect to phosphane exchange (Table 4) cannot arise from the naïve argument of electron donors increasing the electron density available at phosphorus as calculated charges show increased positive charge at phosphorus for **1b** and **1c** relative to **1e**. Instead, the P-N bonding results in a rehybridisation at phosphorus and a change in the nature of the donor orbital. Ultimately, the calculated P-P bond order is (albeit slightly) lower for the internally coordinated salts than for **2e**, and the P-P bonds longer. The increased stability likely therefore arises from the change in degree of P-N electrostatic interaction, whilst the decrease in covalency argues for weaker, more reactive P-P bonds.

**Figure 3:** Variable temperature <sup>31</sup>P NMR showing the dynamic exchange between [2c]BAR<sup>F</sup>, [2e]BAR<sup>F</sup>, and the free phosphanes.

Phosphane-phosphenium systems are highly susceptible to nucleophilic attack by stronger donor species,<sup>23,24</sup> resulting in phosphane exchange. With the above in mind, we reacted one equivalent of Ph<sub>3</sub>P with [2c]BAR<sup>F</sup> to obtain experimental support for the calculated stabilities. This gave immediate formation of a dynamic mixture as seen by <sup>31</sup>P NMR (Figure 3). Immediately upon addition of Ph<sub>3</sub>P, the <sup>1</sup>J<sub>P-P</sub> coupling was lost and the broad peaks seen in the <sup>31</sup>P NMR did not obviously correspond to any free components of the equilibrium. On cooling, however, the peaks sharpened and resolved into free **1c** and [2e]BAR<sup>F</sup> as predicted by the computed relative free energies. This was repeated with [2b]BAR<sup>F</sup> formed *in situ* with the same result; likewise an identical spectrum was formed by the addition of **1b** to [2e]BAR<sup>F</sup>. From these combined computational and experimental results, we must conclude that the increased steric bulk of the *ortho* substituent coupled with rehybridisation at phosphorus counterintuitively lead to weaker donor systems for main group Lewis acids. This is in direct contrast to their behaviour as chelates in which the change in hybridisation does not occur.

With respect to the neutral complexes **3**, for **3b** and **3e** we were only able to locate a minimum corresponding to a weak interaction between donor and Ph<sub>2</sub>PCl and slight elongation of the P-Cl bond relative to free Ph<sub>2</sub>PCl. For **3c**, no such weak complex could be found, but instead significant P-P bonding character and near scission of the P-Cl bond was found in both Fuzzy bond indices and bond lengths. The formation of the complexes, **3**, is enthalpically favourable in all three cases, though the free energy change is positive for all cases, not

unexpected given the entropic costs of binding. No evidence of adduct formation was seen by variable temperature NMR studies on 1:1 mixtures of **1b** or **1c** and Ph<sub>2</sub>PCI. Repeating the experiment using the softer Ph<sub>2</sub>PI gave dramatically different results. Upon combination with one equivalent of **1b**, the <sup>31</sup>P NMR immediately changes – whilst there are still only two major resonances visible, they are broad and centred at δ35.5 ppm and δ-11.8 ppm and do not correspond to any single known species. On cooling the signals continue to move and at -30 °C, <sup>1</sup>J<sub>P-P</sub> coupling becomes resolved. By -50 °C, the <sup>31</sup>P spectrum is essentially identical to that of [**2b**]BAR<sup>F</sup> and is ascribed to the formation of [**2b**]; on warming the spectra revert to those seen at room temperature. Similar results are observed for the reaction of **1c** and **1e** with Ph<sub>2</sub>PI, which converge upon the spectra for [**2c**]I and [**2e**]I. In no case were any signals which could be attributed to neutral adducts observed. From this we conclude that the barriers to interconversion (not calculated) are in all cases small such that the intermediate is too short lived to be observed on the NMR timescale. This would also explain the anomalous calculated structure of **3c**, indicating narrow, shallow potential wells in the energy surface. The resonances observed are therefore simple weighted averages of the signals of **1**, Ph<sub>2</sub>PI and **2** in fast exchange down to -60 °C.

## Conclusions

Herein we report simultaneous inter- and intra-molecular phosphorus Lewis donor-acceptor complex formation when internally solvated triarylphosphanes react with the soft main group Lewis acid [Ph<sub>2</sub>P]<sup>+</sup>, wherein the donor centre also acts as a Lewis acid. This internal solvation results in higher energy lone pairs at phosphorus, arising from rehybridisation of the phosphorus centre due to hypervalent bonding. This does not translate directly to stronger donor-acceptor bonding in phosphane-phosphenium salts, however, due to competing unfavourable steric interactions and increased positive charge at phosphorus. The existence of this binding mode has implications for the utility of such (conventionally) chelating phosphines in main group cation chemistry.

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## Notes and references

- ‡ Full experimental details are available in the supporting information. X-Ray diffraction data have been deposited at the Cambridge Crystallographic Data Centre (CCDC 1429292)
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