Classification: Physical Sciences (Physics)

Two types of quasi-liquid layers on ice crystals are formed kinetically

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Abstract

Surfaces of ice are covered with thin liquid water layers, called quasi-liquid layers (QLLs), even below their melting point (0 °C), which govern a wide variety of phenomena in nature. We recently found that two types of QLL phases appear that exhibit different morphologies (droplets and thin layers) (Sazaki G. *et al*. (2012) Quasi-liquid layers on ice crystal surfaces are made up of two different phases. *Proc*. *Natl*. *Acad*. *Sci*. *USA* 109: 1052−1055). However, revealing the thermodynamic stabilities of QLLs remains a longstanding elusive problem. Here we show that both types of QLLs are metastable phases that appear only if the water vapor pressure is higher than a certain critical supersaturation. We directly visualized the QLLs on ice crystal surfaces by advanced optical microscopy, which can detect 0.37-nm-thick elementary steps on ice crystal surfaces. At a certain fixed temperature, as the water vapor pressure decreased, thin-layer QLLs first disappeared, and then droplet QLLs vanished next, although elementary steps of ice crystals were still growing. These results clearly demonstrate that both types of QLLs are kinetically formed, not by the melting of ice surfaces, but by the deposition of supersaturated water vapor on ice surfaces. This appears to be the first experimental evidence that supersaturation of water vapor plays a crucially important role in the formation of QLLs.

Significance statement

Thin liquid water layers, so-called "quasi-liquid layers" (QLLs), exist on ice surfaces just below the melting point (0°C). The formation of QLLs governs various important phenomena on the earth, such as weather- and environment-related issues, winter sports, etc. Hence, QLLs have attracted considerable attention in the fields of ice physics, meteorology, crystal growth and surface science. Our molecular-level observation of QLLs by advanced optical microscopy revealed that QLLs are formed kinetically as metastable phases only in supersaturated water vapor. This finding opens new horizons of understanding QLLs, which have been so far discussed thermodynamically only at an equilibrium condition.

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Introduction

Ice is one of the most abundant materials on the Earth, and its surfaces are covered with thin liquid water layers even below their melting point (0 °C) (1-4). Such thin liquid water layers are called “quasi-liquid layers” (QLLs). Since QLLs govern the surface properties of ice just below the melting point, it is well acknowledged that surface melting of ice governs a wide variety of phenomena, such as electrification of thunderclouds (4, 5), regelation (4, 6), frost heave (4, 7), conservation of foods, ice-skating (1, 8), preparation of a snow man (1), and growth of ice crystals (2, 4). Therefore, it is essential to understand the surface melting of ice crystals at the molecular level.

After Michael Faraday proposed the existence of QLLs in 1842 (1), many studies experimentally confirmed the formation of QLLs by various methods (See table S1). All of such studies revealed that the thickness of QLLs significantly increases with increasing temperature. However, such the studies utilized spectroscopy and scattering methods, which can obtain only temporally- and spatially-averaged information, or optical microscopy, which does not have sufficient spatial resolution. Hence, the nature of surface melting has not been fully unlocked. To further understand the dynamic behavior of QLLs, we need to perform real-time and real-space observations of ice crystal surfaces at the molecular level.

Recently, we and Olympus Engineering Co., Ltd. have developed one such technique - namely, laser confocal microscopy combined with differential interference contrast microscopy (LCM-DIM) (9), which can directly visualize the 0.37-nm-thick elementary steps on ice crystal surfaces (10, 11). We found that two types of QLL phases with different morphologies appear (12-14): round liquid-like drops (α-QLLs) and thin liquid-like layers (β-QLLs) emerge, irrespective of the face indices of the ice surfaces. Until our recent studies were reported, it had been believed for many years that only one QLL phase appears in the conventional picture of surface melting (2-4). Hence, our results demonstrate that the conventional picture needs important re-examination.

However, it is still unclear how we can explain the generation of the two types of QLL phases. To approach this issue, we first focused our attention on the thermodynamic stabilities of the two types of QLLs. In this study, we examined a water vapor pressure range in which the two types of QLL phases appeared on ice basal faces by LCM-DIM. As a result, we found that the two types of QLL phases are metastable phases that appear only when the water vapor pressure is higher than a certain critical supersaturation.

Results and Discussion

We prepared an observation chamber (Fig. S1) in which the temperature of sample ice crystals T_{sample} and the partial pressure of water vapor P_{H2O} can be separately controlled. We kept the total pressure in the chamber at atmospheric pressure using nitrogen gas, and changed the degree of supersaturation of water vapor $\sigma=(P_{H2O}-P_e)/P_e$ by changing P_{H2O} (here P_e is the solid−vapor equilibrium pressure). We grew Ih ice single crystals on a cleaved AgI crystal (Fig. S1) at T_{sample}=-15.0 °C and P_{H2O}=585 Pa (σ=13 %) in a nitrogen environment. After the preparation of ice single crystals, we increased T_{sample} to a given temperature. All through this process, we kept the sample ice crystals growing, by carefully changing P_{H2O} and confirming the growth by LCM-DIM observations. Then, we observed the behavior of QLLs on basal faces of sample ice crystals under various P_{H2O} (σ) conditions. LCM-DIM images were processed according to the recipe explained in Fig. S2. Details of the control of T_{sample} and P_{H2O} were explained in our previous publication (15).

To reveal the thermodynamic stabilities of the two types of QLL phases, we investigated the water vapor pressure range in which the two types of QLLs could exist on ice basal faces. After we changed T_{sample} and P_{H2O} from -2.4 °C and 585 Pa (σ=13 %) into -2.0 °C and 585 Pa (σ=13 %), both α- and β-QLL phases appeared on a bare ice basal face. Fig. 1 and Video S1 show the ice basal face after the change. In Fig. 1, white and half-black/white arrowheads represent α- and β-QLLs, respectively. The black arrowhead (Fig. 1A) depicts an elementary step (0.37 nm in height) growing laterally in the direction of the black arrow (10, 11, 15). In this study, the differential interference contrast was adjusted as if the ice crystal surface were illuminated by a light beam slanted from the upper left to the lower-right direction (Fig. S3). Hence, the upper left sides and the lower right sides of convex objects showed brighter and darker contrast compared to a flat crystal surface, respectively. As marked by white arrowheads in Figs. 1B and C, α-QLLs newly appeared on a β-QLL phase. In addition, adjacent β-QLLs coalesced with each other (half-black/white arrowheads in Figs. 1B and C), and then β-QLLs grew in the lateral direction.

Next, we reduced P_{H2O} to 568 Pa (to σ=10 %) while keeping T_{sample} constant at -2.0 °C, and observed how the two types of QLLs behaved by LCM-DIM (Fig. 2 and Video S2). The arrowheads and arrows have the same meaning as those in Fig. 1. Elementary steps (black arrowheads) grew laterally in the direction of the black arrow. However, under this condition, the β-QLL (half-black/white arrowheads) receded gradually. Finally the β-QLL disappeared completely (Figs. 2C and D), although α-QLLs (white arrowheads)

remained stable.

Then, we further reduced P_{H2O} to 532 Pa (to σ=3 %), keeping T_{sample} constant at -2.0 °C, and observed the behavior on the ice basal face. Fig. 3 and Video S3 show the result. The elementary step marked by the black arrowhead in Fig. 3E grew laterally in the black arrow direction, and then reached the position marked by the black arrowhead in Fig. 3F. Hence, we could confirm that P_{H2O} was definitely in a supersaturated condition (σ>0). However, the α-QLL in Fig. 3 (white arrowheads) became gradually smaller as time elapsed. Then 2 hours after Fig. 3A was taken, the α-QLL disappeared completely.

In Figs. 1-3, we examined the critical P_{H2O} above which α- and β-QLLs could appear, keeping T_{sample} constant at -2.0 °C. We performed similar observations at various T_{sample} to determine the pressure-temperature range in which two types of QLLs could grow. We summarize the results in Fig. 4. Blue filled circles and red open rhombuses show the minimum P_{H2O} at which α- and β-QLLs could appear, respectively (blue and red solid lines are guides for eyes). In other words, α- and β-QLLs could exist only under P_{H2O} conditions higher than these respective plots. Black open squares present the solid-vapor equilibrium P_{H2O} that we determined experimentally from the LCM-DIM observations of the growth and recession of elementary steps and crystal edges (15). Solid and dashed black curves in Fig. 4 show the solid-vapor equilibrium P_{H2O} and the liquid-vapor equilibrium P_{H2O}, respectively, obtained in previous studies (16, 17). The dash-dotted black line also represents the solid-liquid equilibrium P_{H2O} obtained from a calculation using the molar volumes of water and ice, the melting point, and the heat of fusion of ice crystals (18). Fig. 4 clearly demonstrates that even just below the melting point, α- and β-QLLs appeared only when P_{H2O} was supersaturated for both ice crystals and liquid water. From this result, we proved, for the first time, that α- and β-QLLs are thermodynamically metastable phases. We also found that β-QLLs could appear only in a higher P_{H2O} range than α-QLLs: β-QLLs are less stable than α-QLLs.

Many researchers so far experimentally examined the formation of QLLs under various conditions using different techniques (for detail see Table S1). Some reported that QLLs were formed with P_{H2O} corresponding to an undersaturated condition, and some others reported the formation of QLLs under an equilibrium or supersaturated condition. Hence, until we performed the present experiments, there was not a unified view about the role of supersaturation. Moreover, as we reported previously (Fig. S3 of reference (15)), it is not easy to accurately evaluate the supersaturation of P_{H2O} in the vicinity of a growing or sublimating ice crystal. In our new work, we evaluated the supersaturation of P_{H2O} by directly observing the growth and recession of elementary steps. Hence, we believe that we have evaluated the degree of supersaturation more accurately than the

other experiments performed so far, and that the results shown in Fig. 4 are substantive.

In the conventional picture of surface melting, it had been thought that only one QLL phase fully covers an ice surface at the solid-vapor equilibrium P_{H2O} (2-4). Lacmann and Stranski first gave a thermodynamic explanation for the wetting of ice crystal surfaces with QLLs (19). Then Kuroda and Lacmann developed this issue into the explanation of morphologies of snowflakes (2). Their theoretical model is based on the competition between the stabilities of a bulk (unstable) and interfaces (stable) of QLLs at the equilibrium P_{H2O} (for details see Fig. S4) (2). Kuroda and Lacmann evaluated the wettability of a QLL on ice crystal surfaces utilizing values of interfacial free energies determined previously (2, 18, 20), which lead them to expect that an ice crystal surface is fully wetted with a QLL (2). Taking into account the contributions of a bulk (unstable) and interfaces (stable) of a QLL, Kuroda and Lacmann assumed that a QLL and an ice can be in equilibrium at the solid-vapor equilibrium P_{H2O}, as schematically shown in Fig. 5A. In addition, taking into account the effects of kinetics, they also presumed that a QLL can exist under both undersaturated and supersaturated P_{H2O} conditions (Fig. 5A).

In Fig. 5B, we also summarized the supersaturation range in which the two types of QLLs appeared in our experiments. In contrast to the conventional picture (Fig. 5A), at equilibrium P_{H2O}, we found that no QLL phase appeared on ice crystal surfaces. This discrepancy was mainly due to the misestimation of free energy of an ice-vapor interface by Kuroda and Lacmann, who estimated the ice-vapor interfacial free energy from the density of dangling bonds on ice crystal surfaces. However, at ultra-high temperatures (just below the melting point), the effects of entropy significantly decrease the interfacial free energy. Hence, they overestimated the ice-vapor interfacial free energy, leading to the complete wetting of ice crystal surfaces with QLLs (Fig. 5A). Since we can directly observe two-dimensional (2D) nucleation on an ice crystal surface by LCM-DIM (10, 11), we plan to accurately evaluate the ice-vapor interfacial free energy from the measurement of 2D nucleation rates in the future.

In addition to the situation for equilibrium P_{H2O}, we found that no QLL phase appeared on ice crystal surfaces under both undersaturated and low supersaturated P_{H2O} conditions (Fig. 5B). QLLs appeared only for P_{H2O} higher than a certain critical supersaturation. This result clearly demonstrates that QLL phases are less stable than ice in any P_{H2O} range. In other words, two types of QLLs are formed kinetically, not by the melting of ice surfaces, but by the deposition of supersaturated water vapor on ice surfaces. To our knowledge, this study shows the first experimental evidence that supersaturation of water vapor plays a crucially important role in the formation of

QLLs.

In addition, the two types of QLLs exhibited different stabilities. As demonstrated in this study, β-QLLs appear at a P_{H2O} significantly higher than α-QLLs: the stability of β-QLLs is notably lower than that of α-QLLs. Nevertheless, β-QLLs show a much flatter shape than α-QLLs, implying that β-QLLs have more favorable interaction (*i.e.* better wettability) with ice surfaces than α-QLLs. To solve this contradiction, we need to obtain more information about the structures of α- and β-QLLs. Hence, we are developing a Raman spectrometer combined with LCM-DIM.

The kinetics of the growth and reduction of α- and β-QLLs is very important. However, to obtain meaningful kinetic data, we need to measure separately the mass transfer processes of water molecules (1) between water vapor and QLLs, (2) between QLLs and ice surfaces (effects of QLLs on the growth of ice crystals) and (3) between α- and β-QLLs. At present, we have no practical idea how to measure the processes (1)-(3). Hence, the elucidation of these processes is a major challenge for the future.

In this study, we performed in-situ LCM-DIM observations on ice basal faces to examine the range of P_{H2O} in which α- and β-QLLs emerged. As a result, we found that α- and β-QLLs are metastable phases that can appear only if P_{H2O} is higher than a critical supersaturation. This result indicates that the two types of QLLs are kinetically formed, not by the melting of ice surfaces, but by the deposition of supersaturated water vapor on ice surfaces. Our findings will provide a key to unlocking the secrets of the various phenomena related with surface melting, from the electrification of thunderclouds to the growth of ice crystals (4).

Material and method

A confocal system (FV300, Olympus Optical Co. Ltd.) was attached to an inverted optical microscope (IX70, Olympus Optical Co. Ltd.), as explained in our recent studies (9, 10). A super luminescent diode (Amonics Ltd., model ASLD68-050-B-FA: 680 nm) was used as a light source for the LCM-DIM observations. The LCM-DIM system (Fig. S1) used in this study included a large number of the improvements for the molecular-level observation of ice crystal surfaces under atmospheric pressure (10).

The observation chamber had upper and lower Cu plates, whose temperatures were separately controlled, using Peltier elements (Fig. S1B). At the center of the upper Cu plate, a cleaved AgI crystal (a kind gift from emeritus professor G. Layton of Northern Arizona University), well-known as an ice nucleating agent, was attached using heat grease. On this AgI crystal, the sample ice crystals were grown at -15°C. To supply water vapor to the sample ice crystals, other ice crystals were grown on the lower Cu plate as a source of water vapor. The temperature of the sample ice crystals T_{sample} was determined by controlling the temperature of the upper Cu plate. Since the volume of the source ice crystals on the lower Cu plate was approximately 10^2 times larger than that of the sample ice crystals, the partial pressure of water vapor P_{H2O} inside the observation chamber was dominated by the equilibrium pressure of the source ice crystals. Hence, P_{H2O} was determined by controlling the temperature of the source ice crystals T_{source} (the temperature of the lower Cu plate). Other details, including the calibration of T_{sample} and P_{H2O}, were reported in our recent study (15).

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Figures legends

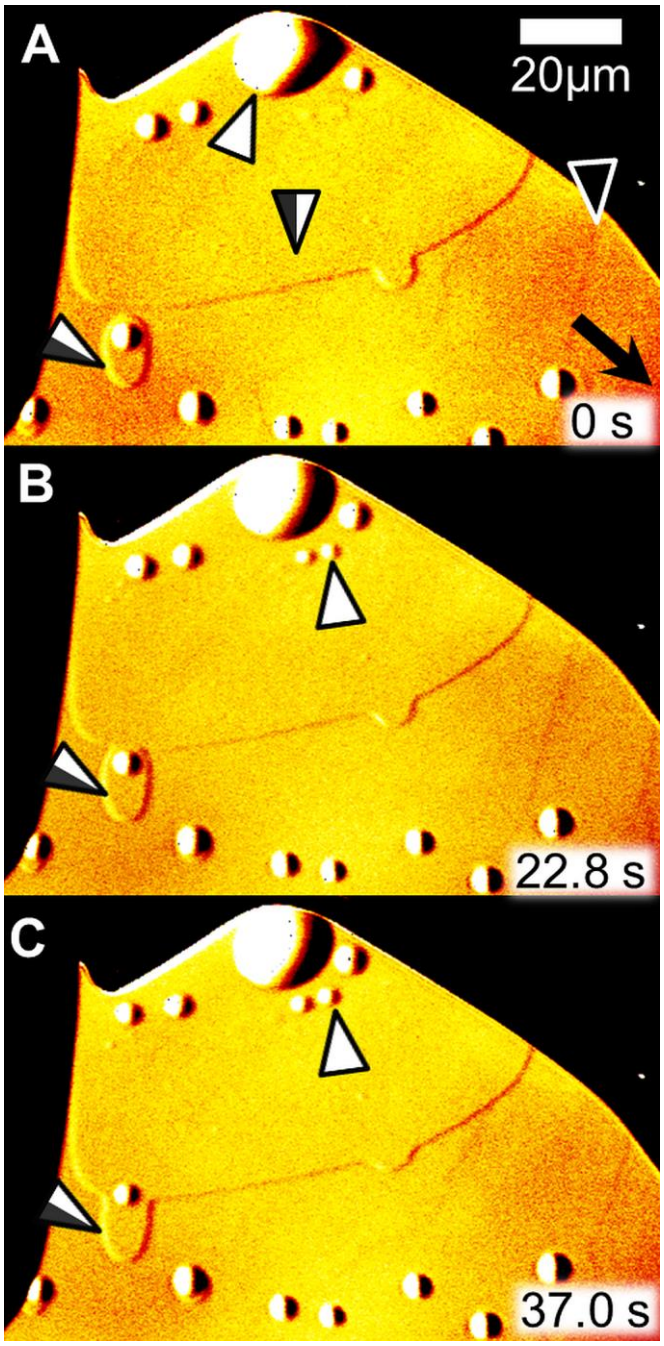

Fig. 1. The appearance of round liquid-like droplets (α-QLLs) and thin liquid-like layers (β-QLLs) on an ice basal face under relatively high supersaturation conditions (P_{H2O}=585 Pa, σ=13 %) at T_{sample}=-2.0 °C. Images B and C were taken 22.8 and 37.0 s after image A, respectively. White and half-black/white arrowheads show α- and β-QLL phases, respectively. A black arrowhead represents an elementary step growing laterally in the black-arrow direction. A movie of the process A-C is available as Video S1.

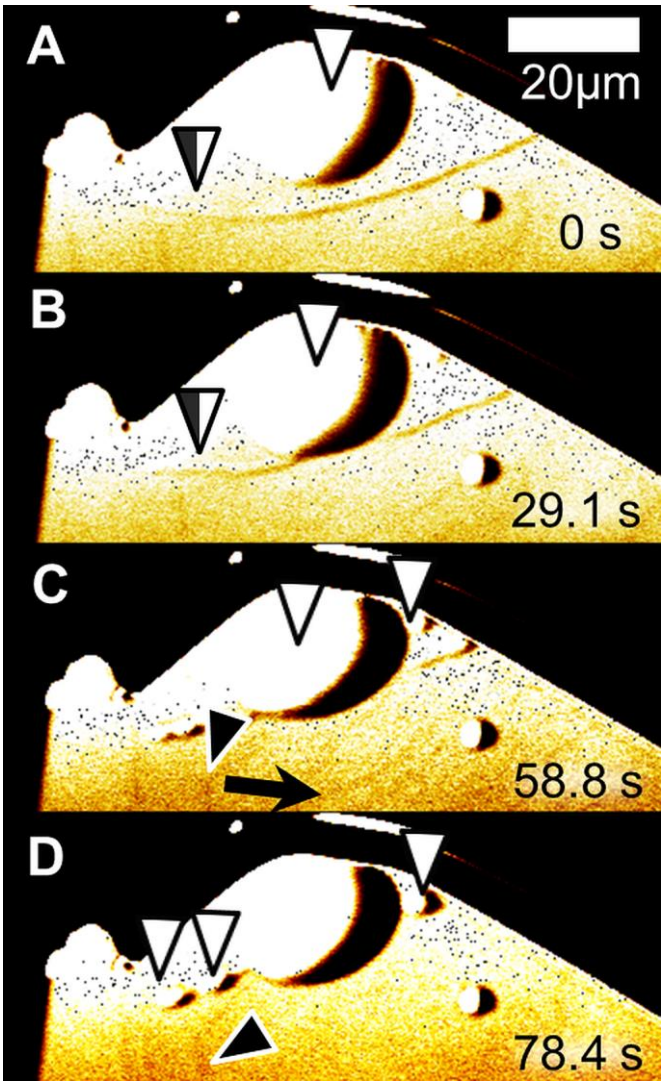

Fig. 2. The disappearance of a thin liquid-like layer (β-QLL) on an ice basal face under relatively medium supersaturation conditions (P_{H2O}=568 Pa, σ=10 %) at T_{sample}=-2.0 °C. After Fig. 1 was taken, P_{H2O} was reduced from 585 to 568 Pa, keeping T_{sample} constant. Images B, C, and D were taken 29.1 s, 58.8 s, and 78.4 s after image A, respectively. White and half-black/white arrowheads show α- and β-QLL phases, respectively. Black arrowheads represent elementary steps growing laterally in the black-arrow direction. A movie of the process A-D is available as Video S2.

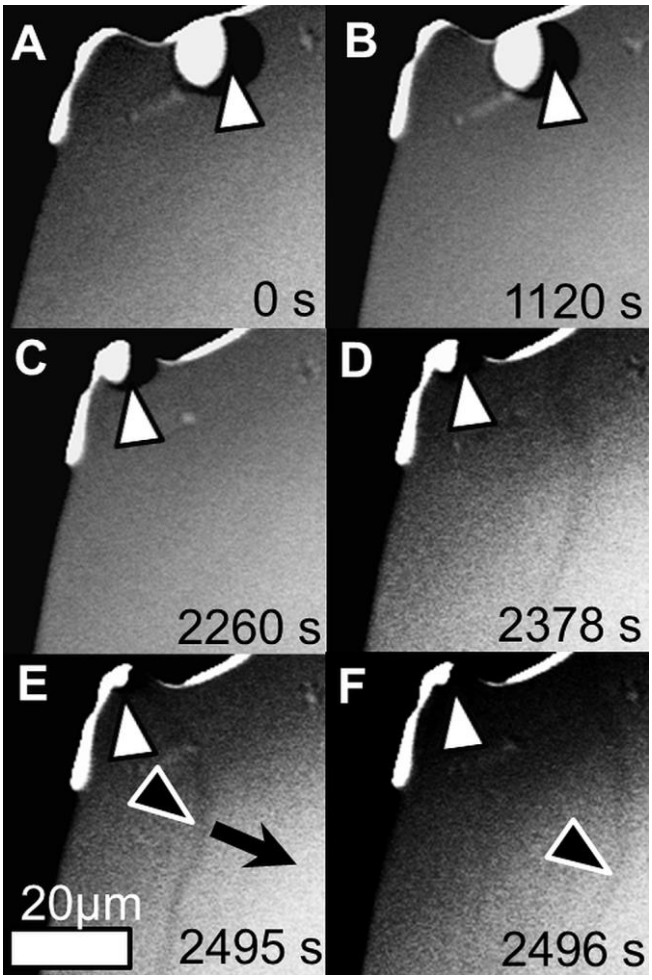

Fig. 3. The disappearance of a round liquid-like droplet (α-QLL) on an ice basal face under relatively low supersaturation conditions (P_{H2O}=532 Pa, σ=3 %) at T_{sample}=-2.0 °C. After Fig. 2 was taken, P_{H2O} was reduced from 568 to 532 Pa, keeping T_{sample} constant. Images B, C, D, E and F were taken 1120 s, 2260 s, 2378 s, 2495 s, and 2496 s after image A, respectively. White arrowheads show α-QLL phases. The elementary step marked by the black arrowhead in E grew laterally in the black arrow direction, and then reached the position marked by the black arrowhead in F. A movie of the process A-F is available as Video S3.

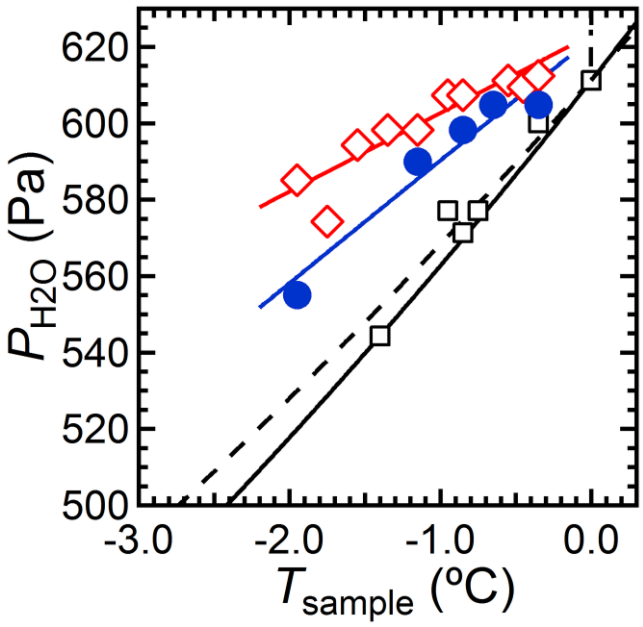

Fig. 4. Pressure-temperature regions in which round liquid-like droplets (α-QLLs) and thin liquid-like layers (β-QLLs) could grow. Blue filled circles and red open rhombuses show the minimum P_{H2O} at which α- and β-QLLs could appear, respectively: α- and β-QLLs could exist only when P_{H2O} is higher than the respective plots. Blue and red solid lines are guides for the eyes. Black open squares presents the solid-vapor equilibrium P_{H2O} that we obtained experimentally from the LCM-DIM observations of the growth and recession of elementary steps and crystal edges (15). Solid and dashed black curves show the solid-vapor equilibrium P_{H2O} and the liquid-vapor equilibrium P_{H2O}, respectively, obtained in the previous studies (16, 17). The dash-dotted black line also represents the solid-liquid equilibrium P_{H2O} obtained from the calculation, using the values shown in ref. 18.

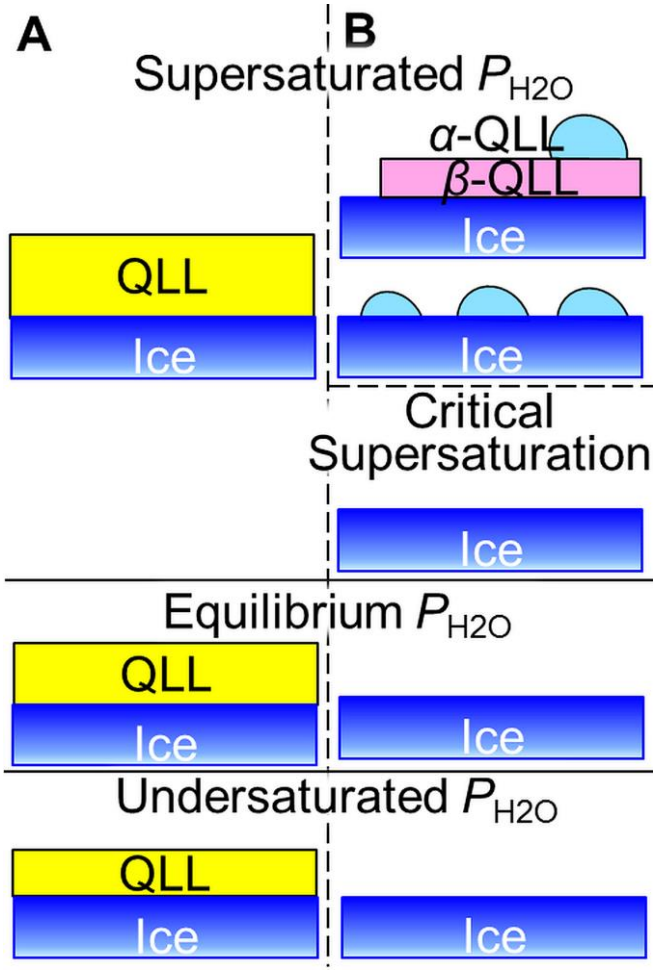

Fig. 5. A comparison between the conventional picture of surface melting (2, 21)(A) and the findings obtained in this study (B). In the conventional picture, one QLL phase appeared irrespective of supersaturation. In contrast, round liquid-like droplets (α-QLLs) and thin liquid-like layers (β-QLLs) appeared only under P_{H2O} conditions higher than the critical saturations.